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Performance of a zeolite modified with *N,N*-dimethyl dehydroabietylamine oxide (DAAO) for adsorption of humic acid assessed in batch and fixed bed columns†

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Factors that affect adsorption of a synthetic humic acid (HA) on a zeolite modified with the surfactant *N,N*-dimethyl dehydroabietylamine oxide (DAAO) (SMZ) were investigated in batch and fixed bed column experiments. Adsorption increased with increasing HA concentrations and contact time, but decreased with increasing pH, temperature, and ionic strength. Adsorption of HA on SMZ was increased by the presence of the cations Ca^{2+} and Mg^{2+} , whereas anions such as NO_3^- , HCO_3^- , SO_4^{2-} , and PO_4^{3-} showed the opposite trend, and competed with HA adsorption. HA adsorption on SMZ was well-fitted by pseudo-second order kinetics, and described by the Langmuir isotherm model. The maximum adsorption capacity in batch experiments calculated from the Langmuir adsorption isotherm was about 126 mg g^{-1} . Thermodynamic calculations showed that HA adsorption on the zeolite with bilayer DAAO coverage was spontaneous and exothermic. Optimum desorption was obtained using 0.1 M NaOH with a recovery of 94%. The HA adsorption capacity of SMZ at the breakthrough point was greatly influenced by bed depth, and could be described by the Thomas model. Adsorption mechanisms are interpreted as involving mainly hydrogen bonding and electrostatic interactions.

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1. Introduction

Humic acid (HA) is an ubiquitous component of ground waters and drinking water sources. However, its presence in water can produce undesirable color and taste, and increase the contents of various pollutants by binding both heavy metals and organic chemicals (e.g., pesticides). More seriously, HA may react with disinfectants (e.g., free chlorine, monochloramine, and ClO_2) during chlorination treatment of drinking water to form carcinogenic disinfection byproducts (DBPs), such as trihalomethanes, haloacetic acids, and haloacetonitriles.¹ Therefore, efficient removal of HA from drinking water supplies or wastewaters is of great importance.

In recent years, various methods have been developed to remove HA from drinking water. These include coagulation/

flocculation, membrane separation, biodegradation, adsorption, ion exchange, and chemical oxidation; among these, adsorption has advantages of low cost, ease of operation, and high efficiency. Various adsorbents have been investigated for the adsorption of HA from water and wastewater, including activated carbon,² resins,³ layered clay minerals,⁴ zeolite,⁵ and chitosan,⁶ but there is still great interest in the development of new more efficient adsorbents.

Zeolites (NZ) are three dimensional aluminosilicate minerals with porous structures, and cationic surfactant-modified NZ have been investigated extensively for the separation and removal of various pollutants from aquatic solutions, including dyes, humic substances,^{5,7} pesticides,⁸ and heavy metals.⁹ Commonly used surfactants for zeolite modification include tetramethylammonium,¹⁰ hexadecyltrimethylammonium (HDTMA),^{11,12} octadecyldimethylbenzyl (ODMBA),^{13,14} and cetylpyridinium bromide.⁵ However, these surfactants are synthetic long-chain molecules with limited biodegradability, and on disposal can result in potentially negative impacts on the environment.^{5,15–17} Thus environmentally-friendly biodegradable surfactants are now attracting more attention.

Rosin obtained from conifer trees is a widely available environmentally friendly material, and consists mainly of resin acids, which are a widely available source of abietic acid.^{18,19} Recently, we reported the synthesis of *N,N*-dimethyl

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dehydroabietylamine oxide (DAAO) and showed that it could modify a zeolite to efficiently remove a dye from aqueous solutions.²⁰ The present work extends these investigations and reports the feasibility of using DAAO-modified zeolites as adsorbents to remove HA from water in batch and fixed bed columns. It considers the effects of various parameters on HA adsorption, including DAAO content, adsorbent dosage, pH, ionic strength, co-existing ions, contact time, and temperature in batch adsorption experiments, and the effects of bed height on HA adsorption and the development of breakthrough curves in fixed column experiments. The experimental data were analyzed by kinetic and isotherm models, and thermodynamic parameters, such as Gibbs free energy change (ΔG_{ads}^0), heat of adsorption (ΔH_{ads}^0), and entropy change (ΔS_{ads}^0) of adsorption were calculated. Possible mechanisms are tentatively proposed for HA adsorption on the DAAO-modified zeolite.

2. Materials and methods

2.1. Materials and reagents

Natural zeolite (NZ) was from the Jinyun County, Zhejiang, China; it had a chemical composition of SiO₂ 74.07%, Al₂O₃ 15.60%, CaO 3.16%, K₂O 2.59%, Na₂O 1.91%, Fe₂O₃ 1.34%, P₂O₅ 0.43%, MgO 0.41%, TiO₂ 0.17%, with a LOI of 8.04%, and an external cation exchangeable capacity (ECEC) of 130 meq kg⁻¹ determined using methylene blue method.^{21,22} Synthetic HA was purchased from Sigma-Aldrich Company (St. Louis, MO, USA). DAAO (purity 93%) was synthesized as described in our previous work,²³ and its physicochemical properties are listed in Table 1. Other chemicals were of the highest purity available, and all solutions were prepared with ultrapure water produced by a Milli-Q system (Advantage A10, Millipore, Billerica, MA, USA).

2.2. Preparation and characterization of the adsorbent

Modified zeolites were prepared by dissolving calculated amounts of DAAO at 60 °C in 250 mL of a water ethanol mixture (50 : 10, v/v) acidified to pH 3.0 with HCl. DAAO-modified zeolites were prepared as described previously.²⁰ In brief, the DAAO stock solution of a 50 mmol L⁻¹ concentration was prepared prior to use in natural zeolite modification. A series of 10 g natural zeolites mixed with different volumes of DAAO solution ranging from 20, 30, 40, 50, 60, 72, and 80 mL, which were named SMZ1, SMZ2, SMZ3, SMZ4, SMZ5, SMZ6

and SMZ7, respectively. The adsorbents were characterized by XRD, FE-SEM and FT-IR spectroscopy before and after adsorption of HA at pH 6.0. XRD patterns were obtained using a Bruker D8 ADVANCE X-ray diffractometer with Cu K α radiation (Siemens D5000 diffractometer, Bruker AXS Inc, Germany), operating at 40 kV and 40 mA. The surface morphologies of solid samples were examined using a JSM-7500F FE-SEM (JEOL Ltd., Japan). Infrared spectra were recorded in KBr pellets on a Nicolet 5700 model FTIR spectrometer (Thermo Nicolet Corporation, USA). Zeta potentials of HA and SMZ were measured using a Zeta sizer 2000 Analyzer (Malvern, Mastersizer 2000 Instruments Co., USA) at an ionic strength of 1 mM NaCl and initial pH ranging from 3.0 to 11.0.

2.3. Batch adsorption experiments

The effects of DAAO load, adsorbent dosage, pH, ionic strength, contact time, and temperature were investigated in batch experiments. Solutions were adjusted to the desired pH by addition of 0.1 mol L⁻¹ HCl or NaOH before the adsorbate. Kinetic studies were carried out at 20 °C for initial HA concentrations in the range 20–100 mg L⁻¹. After predetermined time intervals, aliquots (5.0 mL) were withdrawn and centrifuged immediately at 5000 rpm for 5 min to yield clear supernatant solutions.

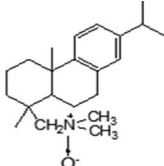
All other adsorption experiments were carried out for 24 h to ensure equilibrium. Adsorption isotherms studies were conducted by varying initial HA concentrations from 10 to 100 mg L⁻¹, by adding a constant amount of adsorbent (0.05 g) at pH 6.0, and then shaking the mixtures at 150 rpm at the specified temperatures. The concentrations of HA remaining in solution were determined by absorption at 254 nm on a UV-Vis spectrometer (GBC Scientific Equipment, Australia). The adsorption amounts and removal efficiency of HA (%) were calculated as follows:

$$q_t = \frac{(C_0 - C_t)V}{m} \quad (1)$$

$$\text{HA removal efficiency\%} = \frac{(C_0 - C_t)}{C_0} \times 100 \quad (2)$$

where C_0 and C_t (mg L⁻¹) are the initial concentrations of HA and at time t , respectively; q_t (mg g⁻¹) is the amount of HA adsorbed at equilibrium; V is the volume of HA solution (L); m is

Table 1 Physicochemical properties of the DAAO used in this study

γ_{cmc} (mN m ⁻¹)	Type	MW (g mol ⁻¹)	Activity (%)	Molecular structure
22–36	Amine oxide	330	93	



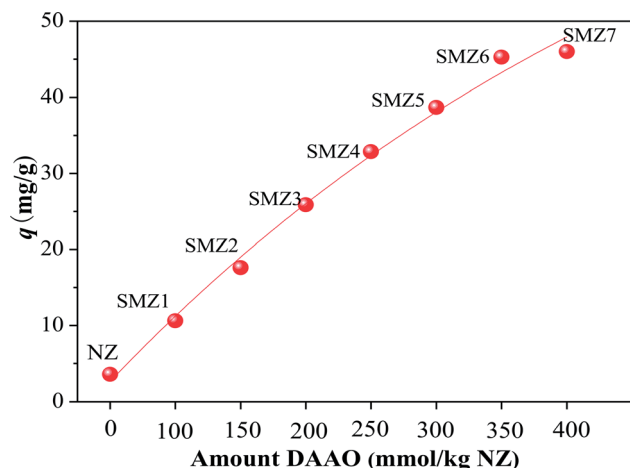


Fig. 1 Effect of DAAO loading on HA adsorption by SMZs. Experimental conditions: $[HA]_0 = 40 \text{ mg L}^{-1}$, [adsorbent dosage] = 1.0 g L^{-1} , contact time = 24 h, pH = 6.0, $T = 20 \pm 1^\circ \text{C}$.

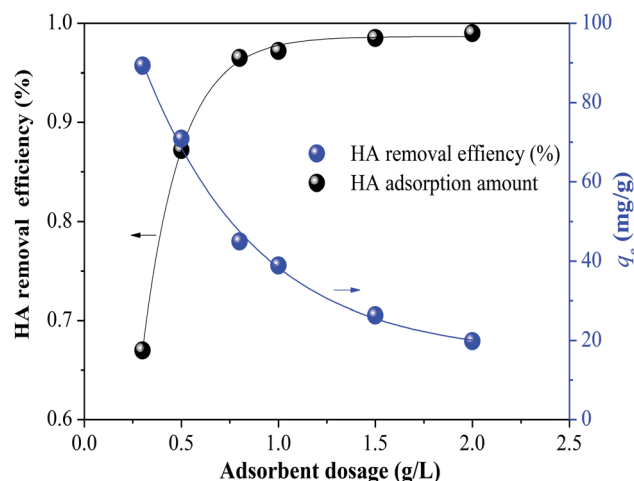


Fig. 2 Effect of dosage of SMZ6 adsorbent on HA adsorption. Experimental conditions: $[HA]_0 = 40 \text{ mg L}^{-1}$, contact time = 24 h, pH = 6.0, $T = 20 \pm 1^\circ \text{C}$.

the mass of adsorbent (g). All adsorption experiments were performed in duplicate.

2.4. Desorption and regeneration studies

In order to better understand the mechanism of HA adsorption on SMZ, desorption measurements were performed in batch experiments with 50 mg SMZ treated with 50 mL HA solution (50 mg L^{-1} , pH 6.0) (derived from Section 2.4) using either 0.1 M NaOH or HCl solutions as desorbing agent. The concentrations of HA desorbed from the adsorbents were determined by absorption at 254 nm as described above. The adsorption–regeneration cycles were repeated four times, and the efficiency of desorption of HA from the adsorbent was calculated by the following equation:

$$\text{Desorption efficiency (\%)} = \frac{q_{\text{de}}}{q_{\text{ad}}} \times 100 \quad (3)$$

where q_{de} and q_{ad} are the HA concentrations (mg L^{-1}) desorbed from SMZ and adsorbed at equilibrium, respectively.

2.5. Fixed-bed column experiments

The effect of bed depth on HA adsorption was studied in glass columns ($\varnothing 1.1 \text{ cm} \times 50 \text{ cm}$) packed to various depths (10, 20 and 30 cm) with SMZ (particle size 0.5–0.8 mm). An initial concentration of 40 mg L^{-1} HA (pH 6.0) was passed through the column at a rate of approximately 4.0 mL min^{-1} and effluent solutions were analyzed for HA at various time intervals.

3. Results and discussion

3.1. HA batch adsorption studies

3.1.1 Effect of amount of DAAO loaded on SMZs on HA adsorption. The amounts of DAAO in SMZ1, SMZ2, SMZ3, SMZ4, SMZ5, SMZ6, and SMZ7 were found to be 90.9, 136.3, 181.5, 222.4, 255.9, 270.9, and 290.2 mmol kg^{-1} (NZ), respectively.²⁰ Thus SMZ1 and SMZ2 had DAAO monolayers, SMZ3,

SMZ4 and SMZ5 had patchy bilayer coverage, and SMZ6 and SMZ7 had DAAO bilayer coverage.²⁰ Results from elemental analyses also showed that DAAO was associated with the zeolite (Table S1†).

The amounts of HA adsorbed on SMZs loaded with different amounts of DAAO are shown in Fig. 1. The original NZ had little affinity for HA, but the HA adsorption efficiency on the SMZs increased with increasing DAAO loading, although there was only a small difference between SMZ6 and SMZ7. Thus maximum adsorption corresponds to complete bilayer DAAO coverage, and SMZ6 was used as the adsorbent in subsequent experiments.

3.1.2 Effect of adsorbent dosage. The effect of adsorbent dosage on HA adsorption by SMZ6 is shown in Fig. 2. There was a major increase in HA adsorption for dosages in the range 0.3 to 0.8 g L^{-1} , but only small increases in the range 1.0 to 2.0 g L^{-1} . Considering efficiency and economy, 0.8 g L^{-1} was chosen as the optimum adsorbent dosage for the following experiments to assess the effects of other factors on the adsorption efficiency.

3.2. Effect of pH

Solution pH is an important parameter that affects the adsorption process. Its effect on HA adsorption by SMZ6 is shown in Fig. 3a, together with the zeta potentials of HA and SMZ6 in the pH range 3.0–11.0. The amount of HA adsorbed by SMZ6 decreased monotonously from 49.4 mg g^{-1} (98.8% removal efficiency) to 7.3 mg g^{-1} (14.7% removal efficiency) as the pH was increased from 3.0 to 11.0. Generally, the pH influences the surface charge, degree of ionization and speciation of both the HA and DAAO, and this result is consistent with the behavior of humic substances associated with soils where they are separated by dissolution in alkali and precipitation with acid. As seen in Fig. 3b, the zeta potential of both SMZ6 and HA decreased progressively with increasing pH. However, SMZ6 has a positively charged surface at low pH, but



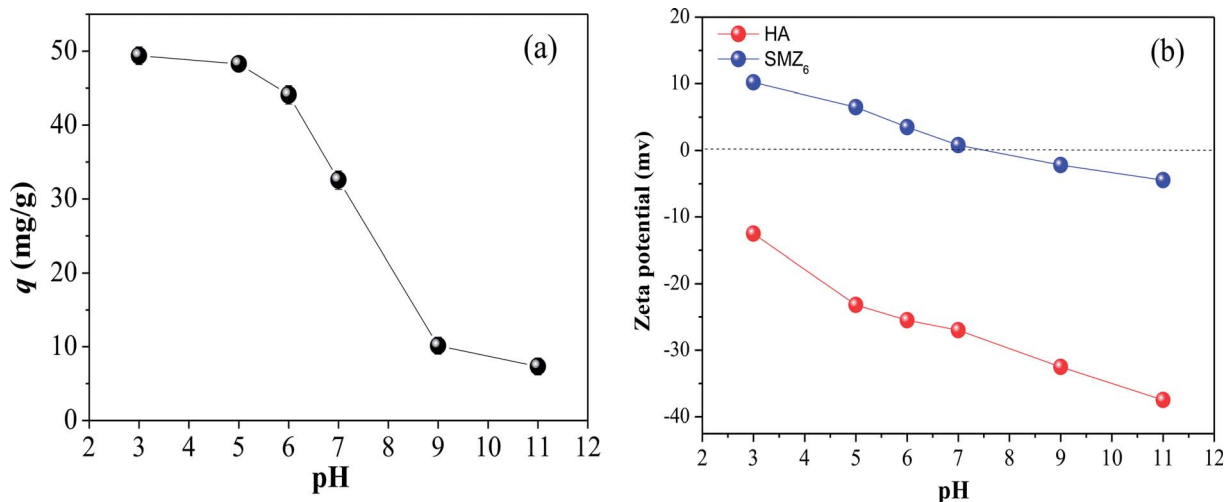


Fig. 3 (a) Effect of pH on HA adsorption on SMZ6, and (b) zeta potentials of HA and SMZ6. Experimental conditions: $[HA]_0 = 40 \text{ mg L}^{-1}$, $SMZ6 = 0.8 \text{ g L}^{-1}$, contact time = 24 h, $pH = 6.0$, $T = 20 \pm 1^\circ \text{C}$.

a negatively charged surface at high pH with the pH_{PZC} at 7.4, whereas HA molecules have negative charges over the entire tested pH range, and their low isoelectric point ($IEP = 1.9$)²⁴ results in deprotonation of carboxyl groups. In the lower pH range ($pH < pH_{PZC}$), protonation of DAAO produces a positive charge center on SMZ (*i.e.* $R-N(CH_3)_2 \rightarrow O + H^+ \leftrightarrow R-N^+(CH_3)_2 \rightarrow OH$, $R = \text{dehydroabietyl base}$), resulting in a strong electrostatic attraction to the $RCOO^-$ groups of HA, and high adsorption of HA on the modified zeolite. In addition, hydrogen bonding may also occur, because DAAO can act as both a hydrogen bonding donor and receptor.²⁰ However, when the pH was increased in the range 7 to 11, HA adsorption decreased sharply, because deprotonation of DAAO results in a loss of electrostatic attraction, and HA adsorption is dependent on hydrogen bonding, and possibly other mechanisms, such as hydrophobic interactions. These conclusions are, therefore, consistent with results reported for HA adsorption on activated

carbon, chitosan hydrobeads, and HTAMB modified attapulgite.^{2,6}

3.3. Effect of ionic strength

The effect of ionic strength on the adsorption of HA on SMZ6 was studied with various concentrations of NaCl. Adsorption decreased linearly with NaCl concentration up to 1.0 M the highest concentration tested (Fig. 4). This result supports the anion exchange mechanism for HA adsorption, because the increasing presence of Cl^- at higher ionic strengths increases competition with HA for the adsorption sites.

3.4. Effect of coexisting ions

Previous studies suggested that oxyanions (*e.g.*, SO_4^{2-} , NO_3^- , HCO_3^- , and PO_4^{3-}) and divalent cations (*e.g.*, Ca^{2+} and Mg^{2+})

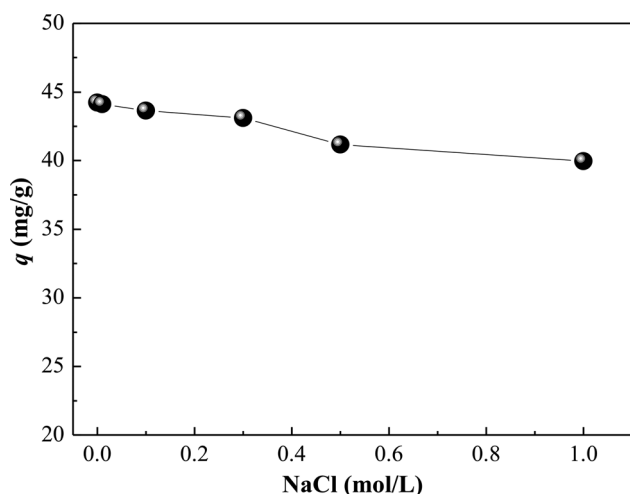


Fig. 4 Effect of ionic strength on HA adsorption on SMZ6. Experimental conditions: $[HA]_0 = 40 \text{ mg L}^{-1}$, $SMZ6 = 0.8 \text{ g L}^{-1}$, contact time = 24 h, $pH = 6.0$, $T = 20 \pm 1^\circ \text{C}$.

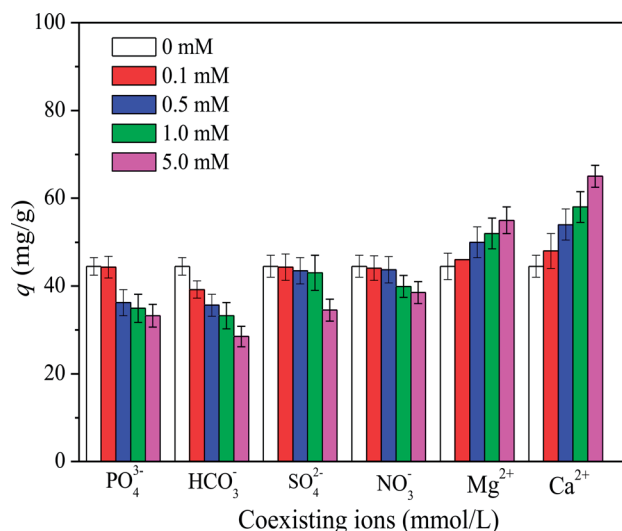
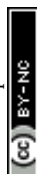


Fig. 5 Effect of co-existing ions on HA adsorption on SMZ6. Experimental conditions: $[HA]_0 = 40 \text{ mg L}^{-1}$, $SMZ6 = 0.8 \text{ g L}^{-1}$, contact time = 24 h, $pH = 6.0$, $T = 20 \pm 1^\circ \text{C}$.



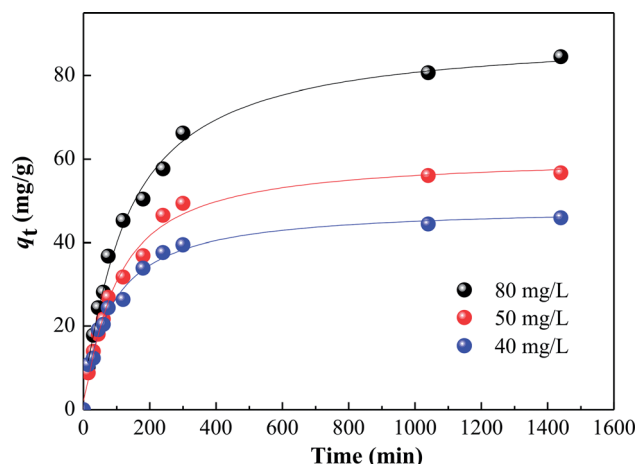


Fig. 6 Kinetics of HA adsorption on SMZ6 at various initial HA concentrations. Experimental conditions: SMZ6 = 0.8 g L⁻¹, pH = 6.0, T = 20 ± 1 °C.

could significantly influence the adsorption of organic matter from natural waters.¹ Fig. 5 shows that HA adsorption on SMZ6 decreased with increasing anion concentrations in the order NO₃⁻ < SO₄²⁻ < HCO₃⁻ < PO₄³⁻. HCO₃⁻ and PO₄³⁻ as Bronsted bases have a strong influence on the HA adsorption, and probably compete with HA for sorption sites on SMZ. Another reason may be attributed to the salting effect on HA adsorption. In contrast, Ca²⁺ and Mg²⁺ had a favorable effect on HA adsorption on SMZ6, probably because of their abilities to combine with anionic groups on both HA and SMZ particles and thereby increase electrostatic attraction as well as bonding directly with the HA and adsorbent. As a result, Ca²⁺ and Mg²⁺ can link the SMZ6 particle with HA, forming an SMZ6-metal-HA complex that can significantly enhance adsorption.^{25,26} Furthermore, direct bonding of the divalent cation Cu²⁺ to both HA and the aluminosilicate mineral was observed by Goodman *et al.*²⁷ in investigations of the adsorption of Cu-HA complexes on montmorillonite and imogolite at pH 7. In that work, the Cu remained bound to the HA when adsorbed on montmorillonite, but was present mainly as the uncomplexed ion on imogolite.

3.5. Adsorption kinetics

The effects of contact time on HA adsorption at various initial HA concentrations (40–80 mg L⁻¹) are shown in Fig. 6. For each concentration, there was fast adsorption in the initial 60 minutes, which then slowed progressively, and became virtually

constant after about 350 minutes. In order to further evaluate the adsorption mechanism, particularly the potential rate-controlling step, four kinetic models, namely pseudo-first-order, pseudo-second-order,²⁸ and intra-particle diffusion²⁹ were used to analyze the data in Fig. 6. From the linear forms of the first three of these models, equations can be written as follows:

First-order kinetic equation:

$$\log(q_e - q_t) = \log q_e - \frac{k_1}{2.303} t \quad (4)$$

Second-order kinetic equation:

$$\frac{t}{q_t} = \frac{1}{k_2 q_e^2} + \frac{t}{q_e} \quad (5)$$

Intraparticle diffusion equation:

$$q_t = k_p t^{0.5} + C \quad (6)$$

where q_t and q_e are the amounts of HA adsorbed (mg g⁻¹) at time t and at equilibrium time, respectively; k_1 (min⁻¹) and k_2 (g mg⁻¹ min⁻¹) are rate constants, respectively. $t^{0.5}$ is the square root of the time and k_p (mg g⁻¹ min^{0.5}) is the intra-particle diffusion rate constant. C is the intercept, which represents the thickness of the boundary layer.

The various kinetic parameters derived from these models are listed in Table 2, which shows that the experimental data fit well to the second-order equation ($R^2 > 0.99$), and the calculated q_e value is in good agreement with the experimental data. However, the intra-particle diffusion model does not fit the whole experimental adsorption data, which can be divided into two or three phases (Fig. 7), corresponding to external surface adsorption followed by internal surface adsorption. This result is consistent with previous findings.³⁰

The adsorption of organic pollutants by porous adsorbents is controlled by external, interfacial, and internal diffusion. In Fig. 8, the kinetic data are also fitted to a simplified homogeneous surface diffusion model (HSDM).³¹

$$\frac{q_t}{q_e} = 1 - \frac{6}{\pi^2} \exp\left(\frac{-D_s \pi^2 t}{R^2}\right) \quad (7)$$

where D_s is the intra-particle diffusion coefficient (cm² s⁻¹), and R is the average adsorbent particle radius. However, the D_s values in the range 6.59–8.81 × 10⁻⁷ cm² s⁻¹ are larger than

Table 2 Kinetic parameters for the adsorption of HA onto SMZ6

[C ₀] (mg L ⁻¹)	q _e (mg g ⁻¹)	Pseudo-first-order			Pseudo-second-order				Intra-particle diffusion		
		k ₁ × 10 ⁻³ (min ⁻¹)	q _{e cal} (mg g ⁻¹)	R ²	k ₂ × 10 ⁻⁴ (g mg ⁻¹ min ⁻¹)	q _{e cal} (mg g ⁻¹)	H	R ²	k _i (mg h ^{0.5} g ⁻¹)	C	R ²
40	45.89	6.2	38.91	0.981	3.2	47.61	1.07	0.997	1.13	11.16	0.789
50	58.66	6.4	53.23	0.986	2.5	61.72	1.83	0.999	1.507	12.5	0.815
80	84.50	2.7	63.84	0.948	0.857	91.74	0.75	0.999	2.239	11.92	0.879



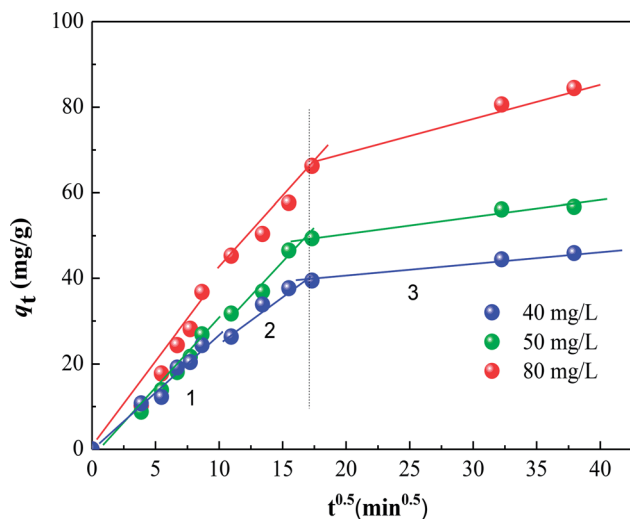


Fig. 7 Intra-particle diffusion model fits to the data for the adsorption of different initial HA concentrations on SMZ6.

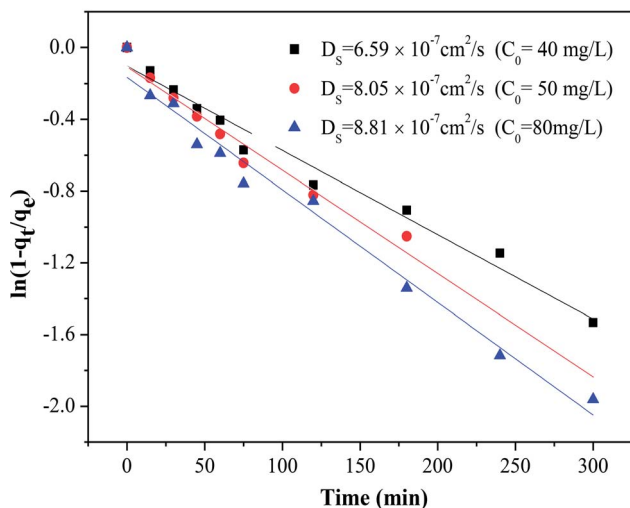


Fig. 8 Intra-particle diffusion model plot for the adsorption of HA on SMZ6.

those reported previously.³² Thus film diffusion may represent the rate-limiting process for HA adsorption on SMZ6. It is postulated that HA molecules are rapidly adsorbed on the external surface of the SMZ6, and that after saturation of the surface, they slowly penetrate the SMZ by intra-particle diffusion through pore and interior surface diffusion until equilibrium is attained. Consequently, the adsorption of HA from aqueous solution using SMZ is a complex process, involving both boundary layer and intra-particle diffusion.

3.6. Adsorption isotherm and thermodynamics

The adsorption isotherms for HA on SMZ6 at 20–40 °C are shown in Fig. 9. The amount of adsorbed HA decreased with increasing temperature, indicating that lower temperatures

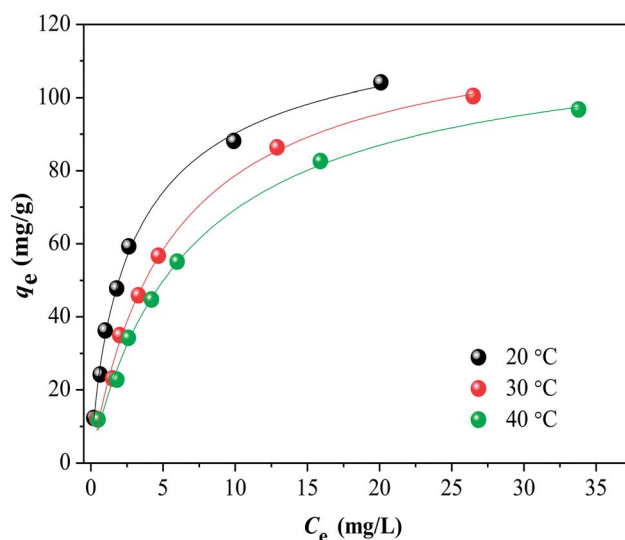


Fig. 9 Isotherms for adsorption of HA on SMZ6 at 20, 30, 40 °C. Experimental conditions: SMZ6 = 0.8 g L⁻¹, pH = 6.0, $T = 20 \pm 1$ °C.

favor the adsorption process. To further understand the adsorption mechanism, the experimental data were analyzed by the Langmuir, Freundlich and Temkin adsorption isotherm models.

Langmuir model:³³

$$q_e = \frac{q_m K_L C_e}{1 + K_L C_e} \quad (8)$$

Freundlich model:³⁴

$$q_e = K_F C_e^{1/n} \quad (9)$$

Temkin model:³⁵

$$q_e = B \ln A_T + B \ln C_e \quad (10)$$

where K_L and q_m are constants related to affinity of the binding sites (L mg⁻¹) and adsorption capacity (mg g⁻¹), respectively; K_F and $1/n$ are the Freundlich constants related to surface heterogeneity of the adsorbent; A and B are the isotherm equilibrium binding constant (L g⁻¹) and isotherm constants in the Temkin model.

The parameters calculated for these isotherms are listed in Table S2.† These show that the data are best fitted by the Langmuir model ($R^2 = 0.99$), and thus HA is adsorbed as a monolayer. The theoretical capacities for HA adsorption on SMZ6 at 20, 30 and 40 °C are 126.6, 123.1 and 119.8 mg g⁻¹, respectively, and are comparable to or greater than those that have been published for other adsorbents (Table 3).

The negative effect of increasing temperature on HA adsorption indicates that the adsorption process is exothermic. The thermodynamic parameters, Gibb's free energy ($\Delta G_{\text{ads}}^\circ$), entropy ($\Delta S_{\text{ads}}^\circ$) and enthalpy ($\Delta H_{\text{ads}}^\circ$) changes for adsorption of HA onto SMZ6 were calculated using the following equations.⁴²



Table 3 Thermodynamic parameters for adsorption of HA onto SMZ6

Adsorbent	q_m (mg g ⁻¹)	T (°C)	pH	Reference
Acid-activated Greek bentonite	12.06	338	3.0	36
Aminopropyl functionalized SBA-15	117.6	25	7.0	37
Bi-functional resin JN-10	105.31	30	11	3
Cetylpyridinium bromide modified zeolite	92.0	25	7.5	5
Chitosan treated granular activated carbon	71.4	25	4.0	30
Crosslinked chitosan-epichlorohydrin beads	44.84	27 ± 1	6.0	38
(HTAB)-modified zeolite synthesized from coal fly ash	126.6	20	3.5	39
Magnetic mesoporous silica composite microspheres (MS-PESI)	128.64	25	6.96	40
Mg/Fe layered double hydroxide	76.70	RT	7.0	41
Polyaniline/attapulgite composite	61.35	25	5.0	24
SMZ6	126.6	20	5.0	This study

$$\Delta G_{\text{ads}}^{\circ} = -RT \ln K_d \quad (11)$$

$$\ln K_d = \frac{\Delta S_{\text{ads}}^{\circ}}{R} - \frac{\Delta H_{\text{ads}}^{\circ}}{RT} \quad (12)$$

where K_d is the adsorption distribution coefficient ($K_d = C_{\text{ad}}/C_e$), C_{ad} is the concentration of solute on the SMZ6 at equilibrium, and R is the universal gas constant (8.314 J mol⁻¹ K⁻¹). The $\Delta H_{\text{ads}}^{\circ}$ and $\Delta S_{\text{ads}}^{\circ}$ parameters were calculated from the slope and intercept of a plot of $\ln K_d$ versus $1/T$ (Fig. S1†), and the calculated thermodynamic parameters are summarized in Table S3.†

The negative values of $\Delta G_{\text{ads}}^{\circ}$ (−5.52 to −6.95 kJ mol⁻¹) at all temperatures indicate the thermodynamic feasibility and spontaneous nature of the adsorption process. Generally, values of $\Delta G_{\text{ads}}^{\circ}$ for physisorption are between −20 and 0 kJ mol⁻¹, and for chemisorption between −400 and −80 kJ mol⁻¹.⁴³ The negative value of $\Delta H_{\text{ads}}^{\circ}$ (−27.93 kJ mol⁻¹) indicates that the adsorption of HA on SMZ6 is an exothermic process. The negative value for $\Delta S_{\text{ads}}^{\circ}$ (−71.78 kJ mol⁻¹) indicates a decrease in randomness at the solid/solution interface during the adsorption process of HA on SMZ6.

3.7. Desorption and regeneration studies

Desorption studies can contribute to the elucidation of adsorption mechanisms, as well as determining the optimum conditions for recovery of the adsorbent. Two different desorption agents, namely 0.1 M NaOH and HCl solutions, were investigated for recovery of HA from the adsorbent. Removal of HA from SMZ6 was favored by high pH, and good desorption efficiency was obtained with 0.1 M NaOH solution. Additionally, the reusability of the adsorbent was tested over four adsorption–desorption cycles (Fig. 10). There was a small progressive decrease in the amount of HA adsorbed after each cycle, but the regenerated adsorbent still possessed about 75% of its original HA adsorption capacity after four cycles, and thus could be used repeatedly.

3.8. Column adsorption performance and models

In addition to the experiments described above, HA adsorption on SMZ6 was also evaluated in fixed bed columns. The effect of bed depth on the breakthrough is shown in Fig. 11 for plots of C_t/C_0 against throughput volume for bed depths of 10, 20, and 30 cm. Generally, the breakthrough point is defined as

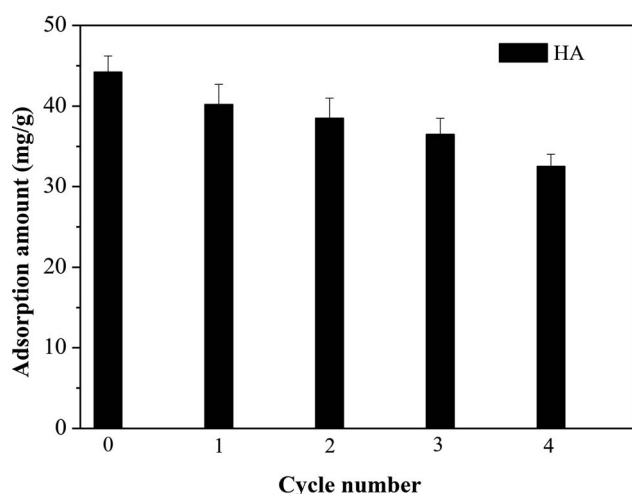


Fig. 10 Amounts of HA adsorbed by SMZ6 after repeated adsorption–desorption cycles. Experimental conditions: $[HA]_0 = 40$ mg L⁻¹, SMZ6 = 0.8 g L⁻¹, pH = 6.0, $T = 20$ °C.

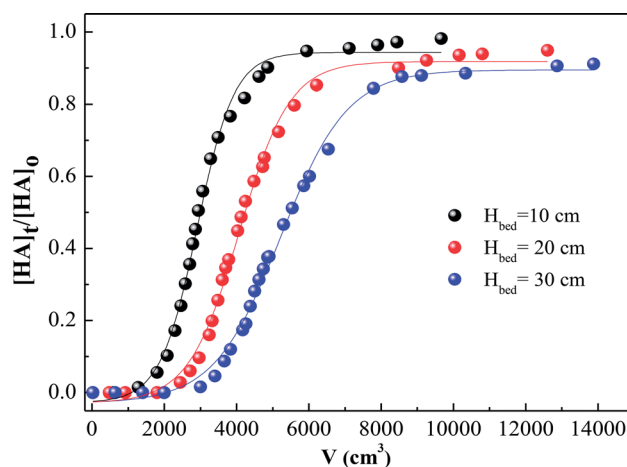


Fig. 11 Breakthrough curves for HA adsorption by SMZ6 fixed columns with different bed depths. Experimental conditions: flow rate = 4.0 mL min⁻¹, $[HA]_0 = 40$ mg L⁻¹, pH = 6.0, $T = 20 \pm 1$ °C.



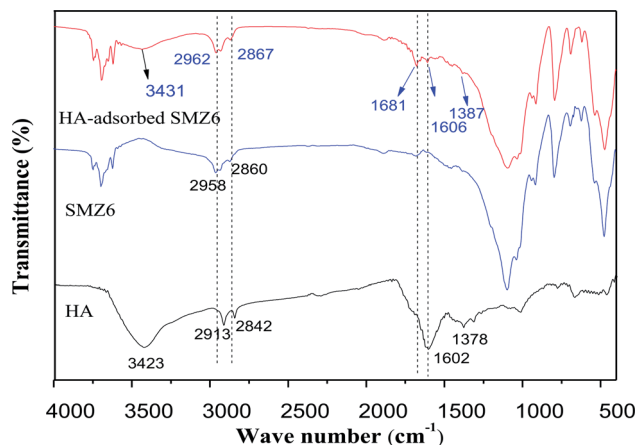


Fig. 12 FTIR spectra of HA, SMZ6 and HA-adsorbed on SMZ6.

the time when the effluent concentration reaches a percentage of the influent concentration (C_0), which is considered unacceptable, e.g. 10% ($C/C_0 = 0.1$). The results in Fig. 11 indicate that an increase in bed depth increases the breakthrough volume and hence the breakthrough time, and results in longer service time. This is mainly due to the increase in the amount of adsorbent packed in the columns, and consequently the number of binding sites with increasing bed depth. For $C/C_0 = 0.1$, the numbers of bed volumes which passed through the adsorbent were, 300, 560 and 620 BV for 10, 20 and 30 cm, respectively. The Thomas,⁴⁴ Yoon–Nelson,⁴⁵ and Adams–Bohart models⁴⁶ were used to analyze the adsorption behavior of HA in this study. The equations and parameters are listed in Text S1.† The breakthrough behavior for the different bed depths is successfully described by the Thomas model ($R^2 = 0.876$ – 0.959), and the dynamic model parameters are summarized in Table S4.†

The bed depth service time (BDST) model based on the theories used for developing the Adams–Bohart model were also used to verify the linear relationship between the bed depth (Z) and saturation time (t_s) of the column.⁴⁷

$$t_s = \frac{N_0}{C_0 V} Z - \frac{1}{C_0 K_a} \ln \left(\frac{C_0}{C_i} - 1 \right) \quad (13)$$

where K_a represents the rate constant in BDST model (min L mg^{-1}), and the other parameters have the same meanings as

described above. The parameters N_0 and K_a can be calculated from the slope of the linear plot of t_s versus Z (Fig. S2†), which is listed in Table S5.† Values of N_0 increased and K_a decreased with increasing values of C_i/C_0 .

The lines of t – Z at values of C_i/C_0 0.1 and 0.9 are shown in Fig. S2.† Plots of breakthrough times (t_b) (corresponding to $C_b/C_0 = 0.1$) and exhaust times (corresponding to $C_b/C_0 = 0.9$) for 10, 20 and 30 cm bed depths are shown in Fig. S2.† and the BDST constants determined from the slopes and intercepts of these lines are listed in Table S5.† This figure shows that $t = 17.8Z - 689.1$ for 10% saturation, and $t = 91.5Z - 546.9$. With the values of $[\text{HA}]/[\text{HA}]_0$ increasing, the values of N_0 increased while K_a decreased. The BDST model parameters can be helpful to scale up the process for other flow rates without further experimental run.

3.9. Characterization of adsorbents before and after adsorption of HA

The FTIR spectra of HA, SMZ6, and HA-adsorbed on SMZ6 are shown in Fig. 12. For HA, the positions of the major absorption bands are: 3423 cm^{-1} from H-bonded OH groups, 1647 cm^{-1} from aromatic $\text{C}=\text{C}$, and 1039 cm^{-1} from $\text{C}-\text{O}$ stretching.⁴⁸ The FTIR spectra of SMZ6 and HA-adsorbed SMZ6 showed two pronounced characteristic peaks at 2958 cm^{-1} and 2860 cm^{-1} that correspond to the stretching vibrations of the $-\text{CH}_2$ and $-\text{CH}_3$ groups of DAAO. The peak at 3431 cm^{-1} in the spectrum of HA-adsorbed SMZ6 that is absent from the spectrum of SMZ6 corresponds to O–H of the HA, and confirms adsorption of HA on SMZ6. After HA adsorption, the asymmetric CH_2 -stretching vibration of DAAO molecules shifted from 2962 to 2958 cm^{-1} and 2860 to 2867 cm^{-1} , thus indicating interactions of HA molecules with alkyl chains of DAAO.

The XRD spectrum of NZ showed the presence of the zeolites clinoptilolite and mordenite along with quartz, and a similar pattern was observed for SMZ6 (Fig. S3†). After adsorbing HA, the structure of SMZ6 did not change, although those at 11.31 , 20.98 , and $25.76^\circ 2\theta$ were decreased in intensity by HA adsorption, suggesting that the adsorption of HA on SMZ6 was on the adsorbent surface.

The sizes and shapes of particles of NZ, SMZ, and HA-loaded SMZ were investigated by SEM (Fig. 13). NZ contained fibrillar crystals of 10–20 nm diameter and variable lengths. However, after modification with DAAO, the zeolite crystals could not be

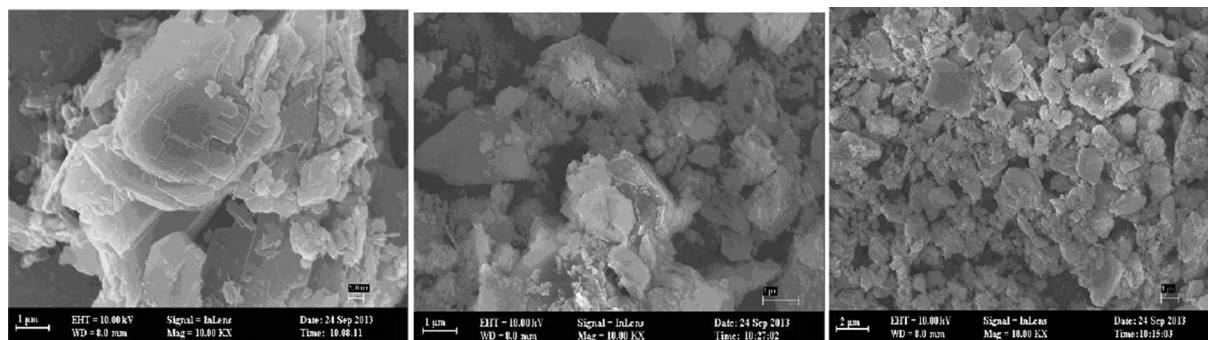
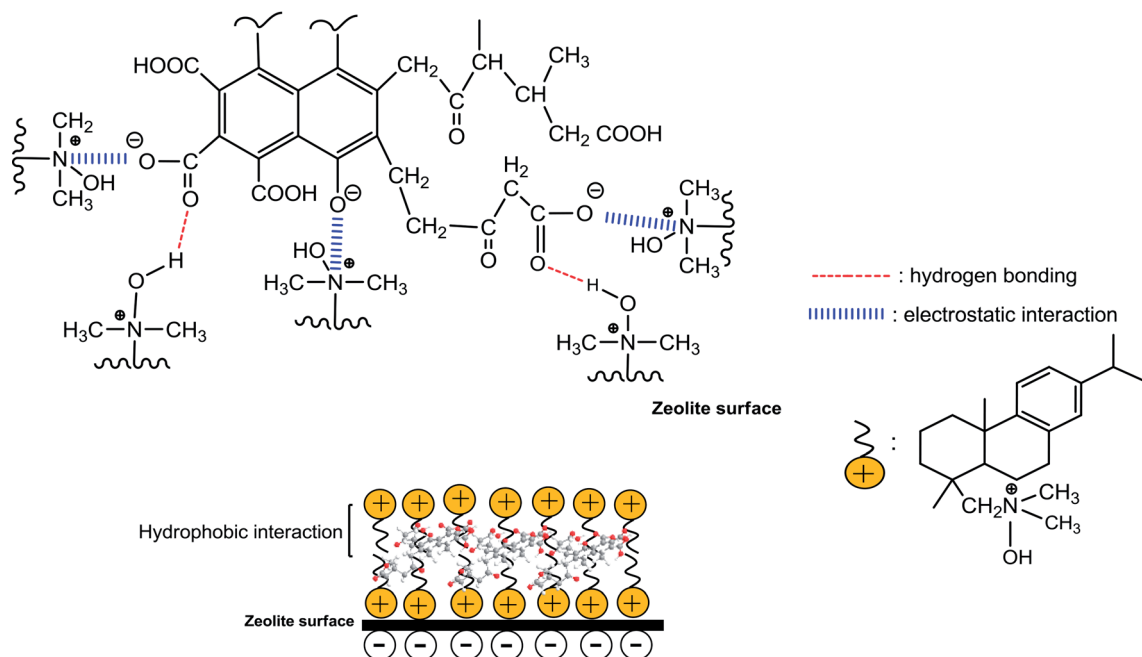


Fig. 13 SEM images of the NZ (a), SMZ6 (b) and HA-SMZ6 (c).



Hypothetical HA model structure (Buffle's model)



Scheme 1 Schematic of interaction mechanisms between HA and SMZ6.

seen clearly (Fig. 13b), thus confirming that their surfaces were covered with DAAO molecules, and further changes in the surface structure occurred after adsorption of HA on SMZ6 (Fig. 13c). BET measurements showed a large decrease in surface area from $65.71 \text{ m}^2 \text{ g}^{-1}$ for the NZ to $3.2 \text{ m}^2 \text{ g}^{-1}$ after modification with DAAO, thus indicating that the surfactant blocked access to the internal pores in the zeolite structure.

3.10. Proposed adsorption mechanism

The following mechanisms have been proposed for the adsorption of HA^{49–51} (i) strong electrostatic interactions between the positively charged outward-pointing head groups of DAAO surfactant bilayers and negatively charged HA molecules in acidic medium; (ii) hydrogen bonding involving the $-\text{N}^+(\text{CH}_3)_2 \rightarrow \text{OH}$ groups of the DAAO molecules, which act as hydrogen donors to carboxylate groups in the HA; (iii) hydrophobic interaction between tails of DAAO monolayers and functional groups of HA molecules in alkaline conditions. These various mechanisms are summarized in Scheme 1. However, mechanism (iii) may be unimportant for adsorption of HA on SMZ6 with DAAO bilayer coverage, since it is removed by 0.1 M NaOH. Furthermore, the results of the present work also suggest an additional mechanism in which the HA molecules are able to slowly diffuse through the DAAO and interact directly with the surface of the zeolite.

4. Conclusions

The HA adsorption capacity, of the natural zeolite was greatly enhanced by modification with the novel surfactant DAAO. The

amounts of HA adsorbed on SMZ decreased with increasing solution pH, temperature, and ionic strength, but were enhanced by the presence of the cations Ca^{2+} and Mg^{2+} . The adsorption process obeys pseudo-second order kinetics, and HA adsorption on SMZ6 was best fitted by the Langmuir isotherm model. The dynamic data were well fitted by the Thomas model. It is concluded that the adsorption mechanisms involve mainly hydrogen bonding and electrostatic interactions. HA was efficiently desorbed from loaded SMZ by alkaline solution, and the regenerated adsorbent retained a high affinity for HA in aqueous solution. Overall this study shows that DAAO modified zeolite can be used as an adsorbent for the removal of HA from aqueous solutions, and can potentially make important contributions to improving drinking water quality by removing HA prior to the disinfection stage.

Conflicts of interest

There are no conflicts to declare.

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