



Cite this: *J. Mater. Chem. A*, 2014, 2, 17867

Correlation between the local scale structure and the electrochemical properties in lithium orthosilicate cathode materials†

Alessandro Mancini,^a Victoria Raissa Barbieri,^a Joerg C. Neufeind,^b Katharine Page^{bc} and Lorenzo Malavasi^{*a}

Lithium metal orthosilicates with general formula Li_2MSiO_4 ($\text{M} = \text{Mn, Fe and Co}$) are among the most attractive new materials as potential high-specific-energy cathodes for lithium batteries. All the members of this family present a rich polymorphism with at least three clearly identified crystal structures of each Li_2MSiO_4 compound. Several theoretical investigations have highlighted that the energy stability of the different polymorphs is very close to each other irrespective of their average crystal structures. At the same time, the calculated and experimental electrochemical performances are again very similar among different polymorphs. By means of neutron total scattering investigation of different polymorphs (monoclinic and orthorhombic) of $\text{Li}_2\text{FeSiO}_4$ and $\text{Li}_2\text{MnSiO}_4$ orthosilicates coupled to Pair Distribution Function (PDF) analysis we showed that, at the local scale, all the polymorphs have the same structure (in particular the structure of the monoclinic polymorph) irrespective of the average structure they possess. This experimental evidence of a strong similarity at the local scale can be correlated with the observed electrochemical similarity (such as the lithium extraction voltages) among the different orthosilicate polymorphs, thus providing an approach to elucidate the relevance of local *versus* long-range structure.

Received 7th August 2014
Accepted 3rd September 2014

DOI: 10.1039/c4ta04063e

www.rsc.org/MaterialsA

Introduction

Lithium batteries are still among the most attractive clean and efficient energy storage systems developed in the last few decades in order to face the well-known problems related to global warming, diminishing fossil-fuel reserves, and increasing demand for portable electronic and grid storage systems.^{1–3} As a matter of fact, the discovery and optimization of high performance materials are critical for next-generation rechargeable batteries.⁴ The Li_2MSiO_4 ($\text{M} = \text{Fe, Mn, Co, and Ni}$) family is very attractive as a positive electrode for lithium batteries due to its improved safety and at least theoretical possibility to deintercalate two Li ions from the structure. In addition, these cathode materials contain abundant, non-toxic and cheap constituent elements. These materials are the object of a very deep and recent experimental and theoretical investigation aimed at improving their properties and performance in Li-batteries.

$\text{Li}_2\text{FeSiO}_4$ and $\text{Li}_2\text{MnSiO}_4$ orthosilicates are characterized by a distorted hexagonal close packing of the oxide ions in which half of the tetrahedral sites are occupied by cations. In this lattice the cation site ordering can vary and the tetrahedra can be distorted thus resulting in a rich and complex polymorphism.^{5–18} The known polymorphs are related to the low- or high-temperature forms of Li_3PO_4 , which are indicated as $\beta\text{-Li}_3\text{PO}_4$ or $\gamma\text{-Li}_3\text{PO}_4$, respectively. Currently, four polymorphs have been reported for $\text{Li}_2\text{FeSiO}_4$ in the literature and crystallized in the $P2_1$, $P2_1/n$, $Pmnb$ and $Pmn2_1$ space groups.^{14,19} Analogously, four crystal structures have been identified for the $\text{Li}_2\text{MnSiO}_4$ orthosilicate with space groups $P2_1/n$, Pn , $Pmnb$ and $Pmn2_1$. The orthorhombic forms ($Pmnb$ and $Pmn2_1$) have a 2D pathway for Li-diffusion while the monoclinic structures ($P2_1$, $P2_1/n$ and Pn) have framework structures with a 3D Li-pathway.

A comparison of the electrochemical properties of the $P2_1$, $Pmnb$ and $Pmn2_1$ $\text{Li}_2\text{FeSiO}_4$ polymorphs, carried out by several groups, showed that the removal of the first Li ion occurs at about 3.1 V during the first charge of the cell, independent of the polymorph.^{19,20} Considering that different average structures correspond to different spatial arrangements of the polyhedra and inter-connections between them, this is a quite unexpected observation. Density Functional Theory (DFT) calculations also showed that the different crystal structures all have similar energies with differences of the order of few kJ mol^{-1} at room temperature.²⁰ In addition, upon delithiation,

^aDepartment of Chemistry and INSTM, University of Pavia, Pavia 27100, Italy. E-mail: lorenzo.malavasi@unipv.it

^bOak Ridge Natl Lab, Chem & Engn Mat Div, Spallat Neutron Source, Oak Ridge, TN 37831, USA

^cLos Alamos Natl Lab, Lujan Neutron Scattering Ctr, Los Alamos, NM 87545, USA

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c4ta04063e



the most stable polymorph is the one characterized by a 3D framework and during this process there is a $2D \rightarrow 3D$ transformation.²⁰

Similar results have been found considering the different $\text{Li}_2\text{MnSiO}_4$ polymorphs which have been the object of several experimental and computational studies as well. As for the previous case ($\text{Li}_2\text{FeSiO}_4$), the thermodynamic stability of the *Pmnb*, *Pmn2₁*, *P2₁/n* and *Pn* structures determined by DFT and atomistic simulation revealed very small energy differences with a slightly higher stability found for the orthorhombic polymorphs.^{14,21} Calculation of the Li-migration pathways for the different structures revealed a strong difference in the activation energies as a function of the space group, passing from 0.65 eV for the *Pn* to 1.13 eV for the *Pmn2₁* structure.²¹ However, the experimental electrochemical behavior of all the $\text{Li}_2\text{MnSiO}_4$ polymorphs is very similar irrespective of the actual crystal structure and framework dimensionality.⁶

This brief overview points out a quite unusual behavior within the orthosilicate cathode materials. Irrespective of the actual average crystal structure, which may vary significantly from orthorhombic to monoclinic, the electrochemical performances and the calculated stabilities as well as Li extraction voltages are very similar. Since these similarities can be hardly correlated with the different average structures found in the Li_2MSiO_4 orthosilicate polymorphs, a possible origin of this behavior may be found in their microscopic structural features. As a matter of fact, the computational studies which found significant differences in the behavior of the various polymorphs (in terms, for example, of ion migration pathways and activation energies) were carried out considering the local structure characteristics. However, the computed relaxed local structures were constrained to be derived from the corresponding average structures.

What is clear from the current literature is that there is a strong discrepancy between the observed experimental results of the electrochemical properties of various orthosilicate polymorphs – *i.e.* all are very close to each other – and the expected differences are based on the average structures. Since such behavior may be closely correlated with the short-range structural features, which have a strong influence on the electrochemical properties (*e.g.*, inter- and intra-polyhedra patterns), in order to shed some light on this aspect, we carried out an experimental investigation of the short-range order in different polymorphs of $\text{Li}_2\text{FeSiO}_4$ and $\text{Li}_2\text{MnSiO}_4$ orthosilicates. This study was carried out by means of neutron total scattering and PDF analysis. Our results indicate that – at the local scale – the structural arrangement of each polymorph corresponds to that of the monoclinic structure and that only at a medium distance range the structure can be properly described considering the relative average structure.

Experimental

The synthesis of the samples investigated in the present work was done as follows. Reagent powders of $\text{LiOH} \cdot \text{H}_2\text{O}$, MnCO_3 , $\text{Fe}(\text{C}_2\text{O}_4) \cdot 2\text{H}_2\text{O}$ and SiO_2 (fumed silica, 7 nm) were purchased from Sigma Aldrich, all of them were of purity >99.9%, and used

without further purifications. Stoichiometric amounts of reactants were mixed in acetone under continuous moderate stirring and N_2 bubbling for 4 hours. Once acetone evaporated, mixed powders were recovered and pelletized using an isostatic oil press at 300 bar for 360 seconds (for *Pmnb* and *P2₁/n-P2₁* polymorphs) or hand-pressed (for *Pmn2₁* polymorphs). Pellets were positioned in a Pt crucible inside a sealed conditioning quartz tube. Argon > 99.9% was then fluxed inside the conditioning tube for 2 hours before starting the annealing process and during all the thermal treatments, in order to prevent Mn^{2+} and Fe^{2+} oxidation. *Pmnb* and monoclinic polymorphs were heated up to 900 °C for 12 h, while *Pmn2₁* was annealed at 700 °C for 8 h. The annealing temperatures were always reached with a heating ramp of 2 °C min⁻¹ while for the cooling process three different paths were followed: cooling ramp of 1 °C min⁻¹ for *Pmnb* polymorphs, cooling ramp of 5 °C min⁻¹ for *Pmn2₁* polymorphs and fast quenching to room temperature for monoclinic polymorphs. Phase purity check was carried out by laboratory XRD (Bruker D8 Advance).

Neutron powder diffraction (NPD) measurements at 300 K were carried out on a NPDP diffractometer at the Lujan Center at the Los Alamos National Laboratory and on a NOMAD diffractometer at the Oak Ridge Laboratories in a cylindrical vanadium tube. Data from an empty container were also collected to subtract the blank. For the NPDP data the corrected total scattering structure function, $S(Q)$, was calculated with the program PDFGETN,²² and the PDF was obtained by Fourier transformation of $S(Q)$. $Q_{\text{max}} = 29.0 \text{ \AA}^{-1}$ was used. For the NOMAD data the $G(r)$ value was extracted by using a program developed by the NOMAD group. Modeling of the experimental neutron PDF data was carried out with the aid of PDFGUI and PDFFit2 software.²³

Results

$\text{Li}_2\text{MnSiO}_4$

Three different polymorphs of the $\text{Li}_2\text{MnSiO}_4$ orthosilicate have been synthesized belonging to the *Pmnb*, *Pmn2₁*, and *P2₁/n* space groups as confirmed by the Rietveld refinement of the neutron diffraction patterns (reported in the ESI†). Table 1 reports the structural parameters obtained from the refinement based on the ND patterns. Our synthesis approach for the sample preparation is a solid-state reaction where the different polymorphs are stabilized as a function of the calcination temperatures and/or cooling ramps. This allowed us to prepare samples with similar morphologies and crystallite dimensions while in the current literature different polymorphs are usually prepared by means of a range of synthesis routes.

Table 1 Lattice parameters for the $\text{Li}_2\text{MnSiO}_4$ orthosilicate polymorphs determined from ND

Sample	Space group	<i>a</i> (Å)	<i>b</i> (Å)	<i>c</i> (Å)	β (°)
$\text{Li}_2\text{MnSiO}_4$ <i>Pmnb</i>	<i>Pmnb</i>	6.3049(2)	10.7430(3)	4.9914(2)	
$\text{Li}_2\text{MnSiO}_4$ <i>Pmn2₁</i>	<i>Pmn2₁</i>	6.3060(2)	5.3683(2)	4.9761(3)	
$\text{Li}_2\text{MnSiO}_4$ <i>P2₁/n</i>	<i>P2₁/n</i>	6.3270(1)	10.9072(2)	5.0691(1)	90.995(1)



Fig. 1a and b present the neutron PDFs for the three $\text{Li}_2\text{MnSiO}_4$ polymorphs discussed in the 1–20 Å range (Fig. 1a) and in the 1–6 Å range (Fig. 1b), respectively.

The simple visual inspection of the neutron PDFs reported in Fig. 1a and b shows that the two $\text{Li}_2\text{MnSiO}_4$ orthorhombic polymorphs share several similarities in terms of peak numbers and shapes while the PDF of the monoclinic structure has a greater number of pairs, which is more evident above *ca.* 10 Å. However, if one looks carefully at Fig. 1b, which is limited to the 1–6 Å range and which encompasses both intra-polyhedra and inter-polyhedra atom pairs, it is possible to observe a strong similarity, in terms of peak number, shape and position, between the PDFs of the three samples considered, irrespective of their average structure.

Fig. 2a–c show the fits of the 1–6 Å range for each $\text{Li}_2\text{MnSiO}_4$ polymorph carried out using as starting parameters those obtained from the Rietveld refinement of the ND patterns and keeping the correct space group for each polymorph.

As can be appreciated, the fit of the data works well for the monoclinic $P2_1/n$ sample while for the other two orthorhombic samples the agreement between the experimental and calculated structures is quite poor (the agreement factors, R_{wp} , are 0.44 ($Pmnb$), 0.26 ($Pmn2_1$), and 0.11 ($P2_1/n$)). This implies that the short-range structure for the two orthorhombic polymorphs should be analyzed by means of another symmetry in order to properly describe the experimental data. This is also in agreement with the data reported in Fig. 2b that suggested a very similar local structure for the three $\text{Li}_2\text{MnSiO}_4$ samples. Accordingly, the fits of the $Pmnb$ and $Pmn2_1$ samples were done in the $P2_1/n$ space group. The lattice parameters for the three samples refined in the $P2_1/n$ symmetry are reported in Table 2

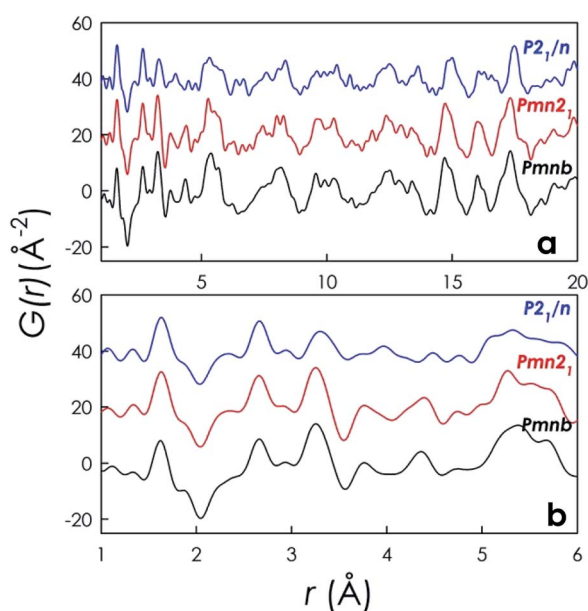


Fig. 1 (a) Neutron PDFs for the three $\text{Li}_2\text{MnSiO}_4$ polymorphs in the 1–20 Å range. Data for the $Pmn2_1$ and $P2_1/n$ structures have been shifted by +20 and +40, respectively. (b) Neutron PDFs for the three $\text{Li}_2\text{MnSiO}_4$ polymorphs in the 1–6 Å range. Data for the $Pmn2_1$ and $P2_1/n$ structures have been shifted by +20 and +40, respectively.

together with the results obtained by fitting with the correct space group.

It is clear that the description of the neutron PDFs with the monoclinic space group for the two orthorhombic polymorphs leads to a very good description of the experimental data. This reduction of symmetry at the local scale is consistent with the similarity in the PDF features discussed above (Fig. 1b). Of course, the monoclinic structure has more refineable parameters than the two orthorhombic structures but we took care to avoid an over-description of the data (for example, we kept fixed the (*x*, *y*, *z*) coordinates of Si and Mn ions in the fits and used one isotropic *a.d.p.* for each type of atom).

In order to further extend the description of the PDF data with respect to the average structure, the data were tested against a medium *r*-range. Fig. 3a–c report the fits in the 11–20 Å range for the three polymorphs carried out with the monoclinic $P2_1/n$ space group used for the 1–6 Å range, while Fig. 3d and e display the same data for the orthorhombic structures fitted within the $Pmnb$ and $Pmn2_1$ symmetry, respectively, *i.e.* with the space group of the average structure.

The R_{wp} for the $Pmnb$ phase (Fig. 3a and d) is lowered from 0.19 to 0.14 if the PDF in the 11–20 Å range is described with the “correct” orthorhombic space group instead of the monoclinic $P2_1/n$. Analogously, for the $Pmn2_1$ sample, R_{wp} changes from 0.24 to 0.17 passing from the $P2_1/n$ to the $Pmn2_1$ space group. This result is quite interesting because it witnesses that, already in a medium range (*i.e.*, 11–20 Å), the description of the experimental PDFs is most correctly done by using the symmetry of the average structure and that the use of a more distorted space group, such as the monoclinic $P2_1/n$ (that more accurately described the short-range order), gives a worse description of the experimental data. This is also a nice proof of the reliability of the fit carried out for the 1–6 Å range.

$\text{Li}_2\text{FeSiO}_4$

The investigation presented above for the $\text{Li}_2\text{MnSiO}_4$ orthosilicates has been carried out in an analogous way for the three polymorphs of the $\text{Li}_2\text{FeSiO}_4$ orthosilicates ($P2_1$, $Pmnb$ and $Pmn2_1$). These have been again prepared by means of a solid state reaction by tuning the temperatures and the heating and cooling ramps in order to get the three different polymorphs. The lattice parameters, as obtained from the refinement of the ND patterns, are reported in Table 3.

Fig. 4 reports the experimental neutron PDFs for the three samples of $\text{Li}_2\text{FeSiO}_4$ investigated. As for $\text{Li}_2\text{MnSiO}_4$, increasing the distance (*r*) leads to a progressive increase of the differences between the three experimental PDFs while, in the short-range, a strong similarity between the data for the three polymorphs is evident irrespective of the average structure.

Fig. 5a–c show the fit in the short range (1–6 Å) of the three experimental PDFs using, as starting parameters, those obtained from the Rietveld refinement of the ND patterns and keeping the space groups of the average structure. The structural results obtained for the fits, together with R_{wp} , are reported in Table 4.



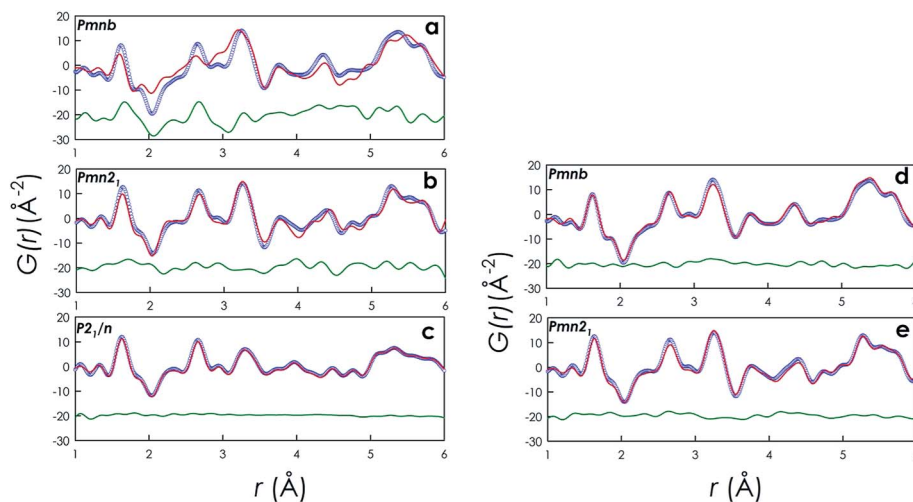


Fig. 2 (a) Fit for the *Pmnb* polymorph with the *Pmnb* symmetry. (b) Fit for the *Pmn2*₁ polymorph with the *Pmn2*₁ symmetry. (c) Fit for the *P2*₁/*n* polymorph with the *P2*₁/*n* symmetry. (d) Fit for the *Pmnb* polymorph with the *P2*₁/*n* symmetry. (e) Fit for the *Pmn2*₁ polymorph with the *P2*₁/*n* symmetry. Legend for all the figures: empty blue circles: experimental data; red lines: fitted structure; green lines: residual. Residuals have been shifted by -20 .

As for the $\text{Li}_2\text{MnSiO}_4$ case, the description of the short-range order for the two orthorhombic polymorphs of $\text{Li}_2\text{FeSiO}_4$ is significantly improved if the data are described using the space group of the most distorted polymorph (*i.e.*, the monoclinic *P2*₁ structure). R_{wp} lowers from 0.236 to 0.114 for the *Pmnb* polymorph and from 0.320 to 0.120 for the *Pmn2*₁ polymorph. Again, in order to avoid an over-description of the experimental data, the atomic positions of the Fe and Si atoms were kept fixed during the refinement and temperature factors were described with only one isotropic factor for both Li and O atoms.

In order to test the validity of the proposed model for a wider r -range, Fig. 6a–e report the refinement of the experimental data in the 11–20 Å range. In particular, Fig. 6a–c report the fits in the 11–20 Å range for the three polymorphs carried out with the monoclinic *P2*₁ space group, *i.e.* the space group that best described the 1–6 Å range, while Fig. 6d and e show the fit results of the two orthorhombic structures carried out with the two orthorhombic space groups (*Pmnb* and *Pmn2*₁).

For $\text{Li}_2\text{FeSiO}_4$, the inspection of the fit results in this range leads to a different conclusion with respect to $\text{Li}_2\text{MnSiO}_4$. For the orthorhombic *Pmnb* phase, the agreement factor is 0.291 when refining the data with the *P2*₁ space group and it lowers to 0.214 if the space group of the average structure is used (*i.e.*, *Pmnb*). This difference is relatively small, and the visual

inspection of the fits reported in Fig. 5a and d does not indicate a huge difference in terms of fit quality between the two structural models. The situation is more clear for the other orthorhombic phase (*Pmn2*₁), where the quality of the fit is good if the data in the 11–20 Å range are again modelled with the *P2*₁ space group instead of *Pmn2*₁ of the average structure (the agreement factor is 0.189 in the first case and 0.389 in the second case). These results indicate that for $\text{Li}_2\text{FeSiO}_4$, the difference between the average and the local structure extends over a wider r -range with respect to $\text{Li}_2\text{MnSiO}_4$. In order to see whether by further extending the analysis it was possible to reconcile the average and local structure, the two orthorhombic samples were tested against the two possible models (*i.e.*, the monoclinic and the corresponding orthorhombic space groups) up to 30 Å.

The results of the fits from 21 to 30 Å for the two orthorhombic polymorphs of $\text{Li}_2\text{FeSiO}_4$ clearly indicate a huge improvement of the results if the average structure space group is used instead of the monoclinic one. The agreement factor lowers from 0.506 to 0.189 for the *Pmnb* polymorphs fitting the data from 21 to 30 Å with the *P2*₁ space group or the *Pmnb* structure, respectively. On the same line behaves the *Pmn2*₁ polymorph, where the description of the local structure with the orthorhombic space group lowers R_{wp} from 0.365 (*P2*₁ space group) to 0.156 (*Pmn2*₁ space group). The

Table 2 Fit results of the neutron PDFs for the $\text{Li}_2\text{MnSiO}_4$ polymorphs (1–6 Å)

Sample	Space group (average)	Space group fit 1–6 Å	a (Å)	b (Å)	c (Å)	β (°)	R_{wp}
$\text{Li}_2\text{MnSiO}_4$	<i>Pmnb</i>	<i>P2</i> ₁ / <i>n</i>	6.4353(6)	10.7173(9)	4.8999(8)	93.74(8)	0.131
$\text{Li}_2\text{MnSiO}_4$	<i>Pmn2</i> ₁	<i>P2</i> ₁ / <i>n</i>	6.4019(7)	10.746(1)	4.9177(5)	90.78(2)	0.145
$\text{Li}_2\text{MnSiO}_4$	<i>P2</i> ₁ / <i>n</i>	<i>P2</i> ₁ / <i>n</i>	6.620(1)	11.040(1)	5.0915(3)	92.305(7)	0.109
$\text{Li}_2\text{MnSiO}_4$	<i>Pmnb</i>	<i>Pmnb</i>	6.2189(7)	11.466(1)	4.9622(4)	90	0.443
$\text{Li}_2\text{MnSiO}_4$	<i>Pmn2</i> ₁	<i>Pmn2</i> ₁	6.1510(5)	5.4927(3)	5.0314(3)	90	0.255



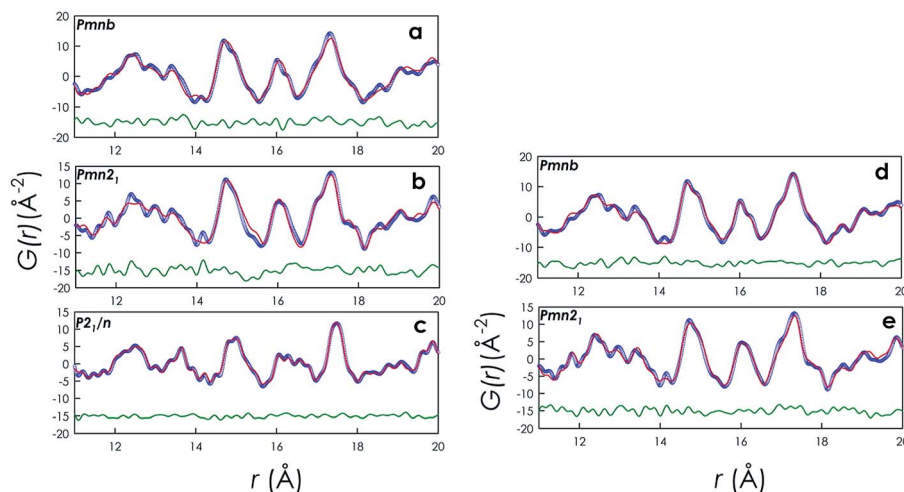


Fig. 3 (a) Fit for the *Pmnb* polymorph with the *P2₁/n* symmetry. (b) Fit for the *Pmn2₁* polymorph with the *P2₁/n* symmetry. (c) Fit for the *P2₁/n* polymorph with the *P2₁/n* symmetry. (d) Fit for the *Pmnb* polymorph with the *Pmnb* symmetry. (e) Fit for the *Pmn2₁* polymorph with the *Pmn2₁* symmetry. Legend for all the figures: empty blue circles: experimental data; red lines: fitted structure; green lines: residual. Residuals have been shifted by -15 .

Table 3 Lattice parameters for the $\text{Li}_2\text{FeSiO}_4$ orthosilicate polymorphs determined from ND

Sample	Space group	<i>a</i> (Å)	<i>b</i> (Å)	<i>c</i> (Å)	β (°)
$\text{Li}_2\text{FeSiO}_4$	<i>Pmnb</i>	6.2660(2)	10.6434(2)	5.0210(1)	
$\text{Li}_2\text{FeSiO}_4$	<i>Pmn2₁</i>	6.3049(2)	10.7430(3)	4.9914(2)	
$\text{Li}_2\text{FeSiO}_4$	<i>P2₁</i>	8.2257(1)	5.0056(3)	8.2142(1)	99.220(4)

corresponding plot of the refinements up to 30 Å is reported in the ESI.†

Discussion and conclusions

Lithium orthosilicates are very attractive materials in the current research as possible Li-battery cathodes even though the proposed two-Li ion exchange process is unlikely to be realized. However, the orthosilicates have also shown problems related to achieving phase purity. On the other hand, the electrochemical and theoretical investigations on lithium orthosilicates have shown strong similarities irrespective of the exact polymorph under investigation and, in addition, it seems that all the starting structures tend to stabilize the same delithiated phase.²⁰ As we discussed in the Introduction section, this seems to be a quite unusual behavior in the field of Li-cathode materials where the electrochemical properties are strongly connected to the structural features, which in turn influence the Li-migration paths. However, it is also true that most of the electrochemical probes are closely affected by the actual local structure of the investigated materials and that the computed electrochemical properties are based on local structure optimization tools.

In this work we have considered in close detail the local structure behavior of three polymorphs (two orthorhombic and one monoclinic) for the two orthosilicates $\text{Li}_2\text{MnSiO}_4$ and $\text{Li}_2\text{FeSiO}_4$. Discussing first the average structure, each polymorph can be described in a reasonable way with the corresponding space group model if neutron diffraction patterns are taken in consideration and analyzed by means of Rietveld refinement. However, the most careful investigations of the structural features of Li orthosilicates that can be found in the current literature point towards the presence of small deviations from the average structures employed to model the X-ray or neutron diffraction data.¹⁰ All these observations were carried

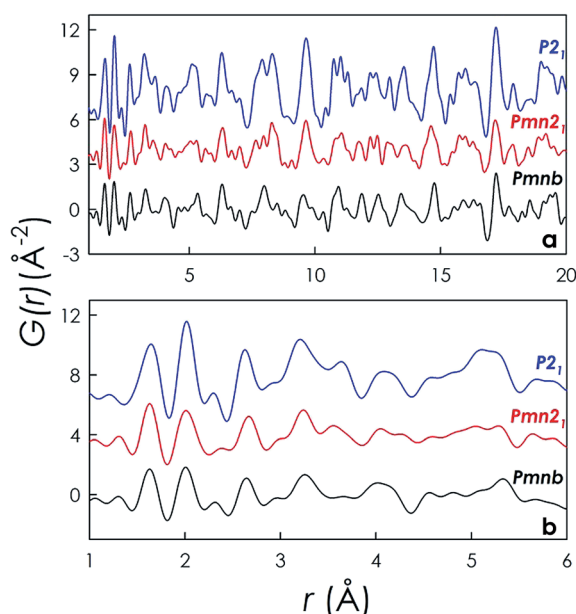


Fig. 4 (a) Neutron PDFs for the three $\text{Li}_2\text{FeSiO}_4$ polymorphs in the 1–20 Å range. Data for the *Pmn2₁* and *P2₁* structures have been shifted by +4 and +8, respectively. (b) Neutron PDFs for the three $\text{Li}_2\text{FeSiO}_4$ polymorphs in the 1–6 Å range. Data for the *Pmn2₁* and *P2₁* structures have been shifted by +4 and +8, respectively.



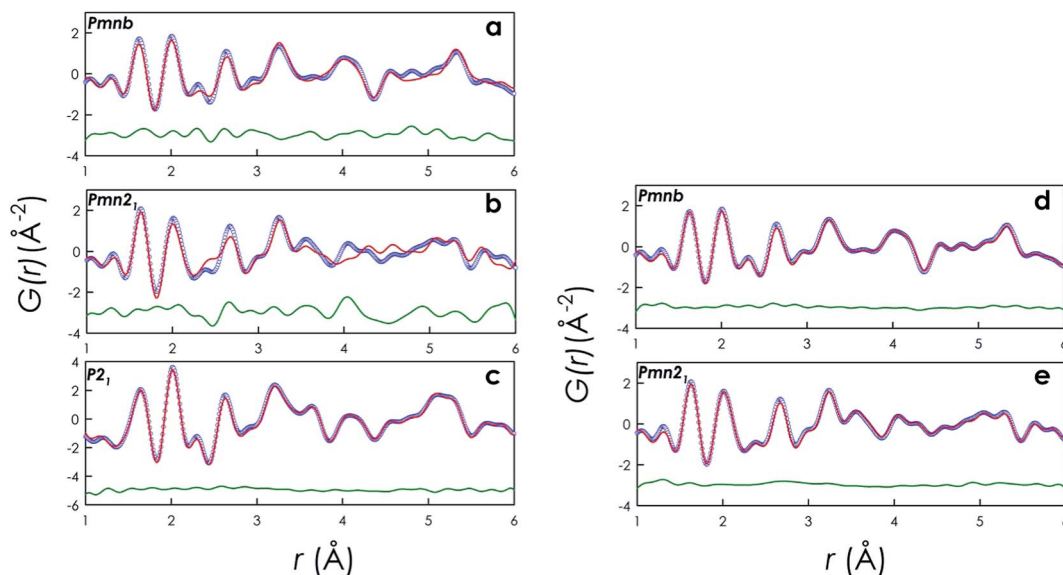


Fig. 5 (a) Fit for the *Pmnb* polymorph with the *Pmnb* symmetry. (b) Fit for the *Pmn2*₁ polymorph with the *Pmn2*₁ symmetry. (c) Fit for the *P2*₁ polymorph with the *P2*₁ symmetry. (d) Fit for the *Pmnb* polymorph with the *P2*₁ symmetry. (e) Fit for the *Pmn2*₁ polymorph with the *P2*₁ symmetry. Legend for all the figures: empty blue circles: experimental data; red lines: fitted structure; green lines: residual. Residuals have been shifted by -5 .

out by means of transmission electron microscopy (TEM) which can provide important structural information at the local scale but which cannot be considered a bulk probe. Deviations from the expected long range is also visible from the presence of modulations in the background of high quality ND patterns and, more clearly, in the inspection of the $S(Q)$ vs. Q data we collected in this experiment.

The analysis of the neutron PDFs for the three different polymorphs of each compound has led to an important observation. Considering the local structure in the range up to 6 Å all the polymorphs of $\text{Li}_2\text{MnSiO}_4$ and $\text{Li}_2\text{FeSiO}_4$ are properly described by making use of a single monoclinic structure irrespective of the average structure. This range roughly corresponds to one unit cell of the silicates. Extending the PDF analysis up to 20 or 30 Å, the different polymorphs are properly described making use of the average structure space groups reported in the literature and used to analyze the ND patterns. From a structural point of view, the use of a monoclinic local structure, also for the orthorhombic samples indicates that, at the local scale, the lithium orthosilicates are more distorted than expected from the average structure. Fig. 7 reports an illustrative graphical representation of the changes occurring at

the structure passing from an orthorhombic (*Pmnb*) to a monoclinic symmetry (*P2*₁/*n*).

The use of a monoclinic local structure leads to a broader bond-length distribution with respect to the orthorhombic symmetry. For example, the Mn(Fe)–O and Si–O polyhedra have three distinct bond lengths (with two equal distances between Si (and Mn(Fe)) and the O-atoms running along the *a*-axis) ranging from *ca.* 1.65 Å to more than 2 Å in the orthorhombic symmetry. The use of monoclinic local symmetry results in four bond lengths within the same distance range (bond lengths for the samples investigated are reported in the ESI†). In addition, the PDF results indicated that, at the local scale, all the polymorphs present two distinct Li-sites with respect to the single Li site in the orthorhombic structures. Finally, the slight deviation of the β angle from 90° induces some rotation of the Si–O and Mn(Fe)–O polyhedra.

From our neutron PDF results it seems quite clear that the increased disorder of the orthorhombic polymorphs in the orthosilicates is mostly confined at the local scale. The common structural features among all the polymorphs at the local scale have important consequences on the electrochemical performance. It is well known that the local

Table 4 Fit results of the neutron PDFs for the $\text{Li}_2\text{FeSiO}_4$ polymorphs (1–6 Å)

Sample	Space group (average)	Space group fit 1–6 Å	<i>a</i> (Å)	<i>b</i> (Å)	<i>c</i> (Å)	β (°)	<i>R</i> _{wp}
$\text{Li}_2\text{FeSiO}_4$	<i>Pmnb</i>	<i>P2</i> ₁	8.294(1)	5.0452(5)	8.3694(4)	102.27(1)	0.114
$\text{Li}_2\text{FeSiO}_4$	<i>Pmn2</i> ₁	<i>P2</i> ₁	8.2912(5)	5.0758(5)	8.2567(1)	100.60	0.120
$\text{Li}_2\text{FeSiO}_4$	<i>P2</i> ₁	<i>P2</i> ₁	8.3195(7)	5.0655(3)	8.3173(3)	102.81(1)	0.091
$\text{Li}_2\text{FeSiO}_4$	<i>Pmnb</i>	<i>Pmnb</i>	6.281(1)	10.571(2)	4.982(1)	90	0.236
$\text{Li}_2\text{FeSiO}_4$	<i>Pmn2</i> ₁	<i>Pmn2</i> ₁	6.5052(5)	4.8237(4)	4.5911(5)	90	0.320



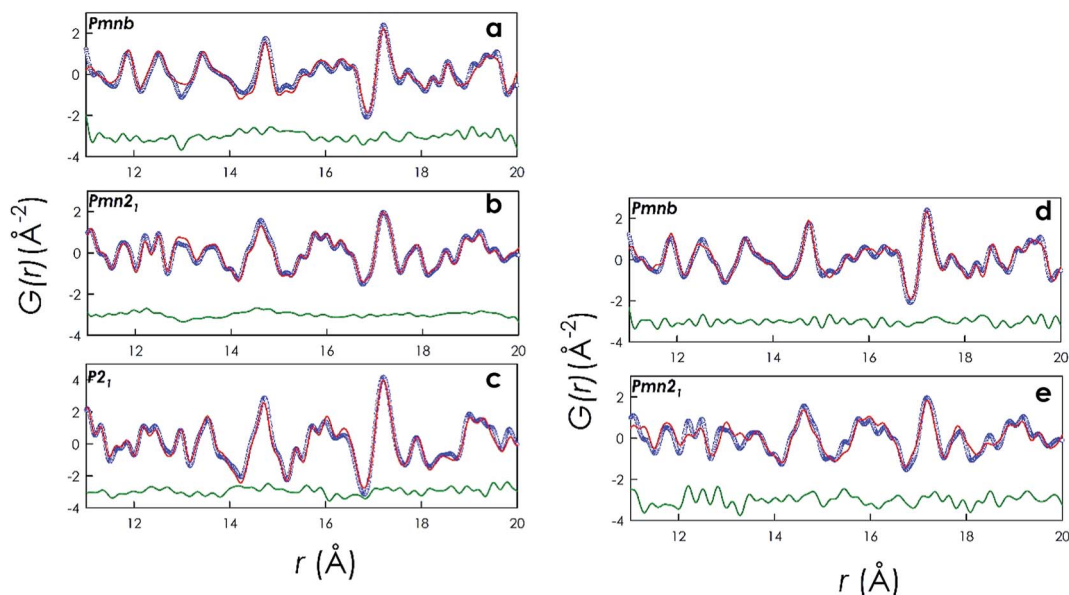


Fig. 6 (a) Fit for the $Pmnb$ polymorph with the $P2_1$ symmetry. (b) Fit for the $Pmn2_1$ polymorph with the $P2_1$ symmetry. (c) Fit for the $P2_1$ polymorph with the $P2_1$ symmetry. (d) Fit for the $Pmnb$ polymorph with the $Pmnb$ symmetry. (e) Fit for the $Pmn2_1$ polymorph with the $Pmn2_1$ symmetry. Legend for all the figures: empty blue circles: experimental data; red lines: fitted structure; green line: residual. Residuals have been shifted by -3 .

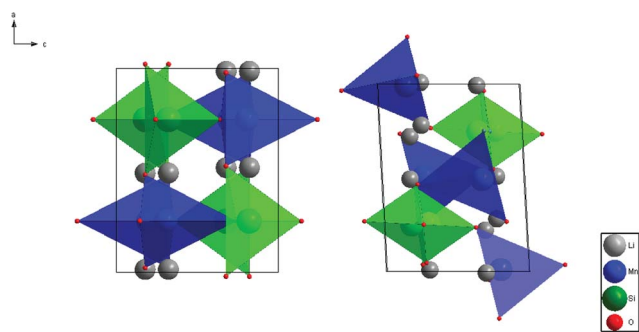


Fig. 7 Arrangement within the unit cell for $\text{Li}_2\text{MnSiO}_4$ in the $Pmnb$ s.g. (a) and in the $P2_1/n$ s.g. (b). Structural data taken from the PDF refinements.

arrangement of Li-ions has the strongest impact on the Li-diffusion and, in turn, on the electrochemical behavior. A recent modelling study has shown that orthorhombic polymorphs of lithium silicates have simpler Li migration paths with respect to the monoclinic structures but that the activation energies for the higher symmetric structures have higher energies with respect to the monoclinic arrangements.²¹ For example, the first jump in the $Pmn2_1$ and $Pmnb$ polymorphs of $\text{Li}_2\text{MnSiO}_4$ has been calculated to be more than 1 eV with distances above 3 Å while in the Pn symmetry E_a is around 0.65 eV with a jump distance of below 3 Å. Second jump energies and distances among the different polymorphs are very similar. The results of this modelling work indirectly corroborate the actual PDF results reported in this work. Most of the electrochemical properties are strongly linked to the Li extraction and in particular to the local

arrangement of Li ions within the orthosilicate structure. The extremely similar experimental extraction voltages suggest that in all the polymorphs the Li-environment is very similar leading to similar electrochemical results, which should not be the case if very different activation energies and local migration paths are present. For example, the modelling results predict that the electrochemical properties of the $Pmn2_1$ polymorph of $\text{Li}_2\text{MnSiO}_4$ should be significantly poorer with respect to the other polymorphs by considering the activation energies and migration paths at the sort range as calculated by using an energy relaxed local structure based on the $Pmn2_1$ symmetry. However, this is not the case, and this polymorph behaves in the same way as the other $\text{Li}_2\text{MnSiO}_4$ polymorphs. The results of the present work indicate as well that the delithiation performance does not depend on the long range crystal structure of the selected polymorph and that the efforts in optimizing the lithium orthosilicate properties should not be directed towards the optimization of highly phase-pure materials since even mixtures of different polymorphs should work well. This is further true in view of the fact that after delithiation all the polymorphs tend to transform to a single stable structure.

Finally, the results reported here also support the evidence that the Li-cathode material performance is mostly dependent upon the local structure of Li-ions within a peculiar structure and that the actual longer range diffusion (being 2D or 3D) has a less important impact on the electrochemical properties. In view of this comprehensive structural investigation of the local scale structure of prototype Li-battery cathode materials (*i.e.*, orthosilicates), it can be argued that the design of improved and optimized new compositions should pay strong attention to the local structural features more than the symmetry on the long range.



Acknowledgements

Funding from the INTSM-RL research project "ATLANTE" is acknowledged. Research conducted at the ORNL's Spallation Neutron Source was sponsored by the Scientific User Facilities Division, Office of Basic Energy Sciences, and US Department of Energy. This work has benefited from the use of the Lujan Center at the Los Alamos Neutron Science Center, funded by the Department of Energy Office of Basic Energy Sciences.

References

- 1 M. S. Whittingham, *MRS Bull.*, 2008, **33**, 411.
- 2 B. Dunn, H. Kamath and J.-M. Tarascon, *Science*, 2011, **334**, 928.
- 3 J. Liu, J.-G. Zhang, Z. Yang, J. P. Lemmon, C. Imhoff, G. L. Graff, L. Li, J. Hu, C. Wang, J. Xiao, G. Xia, V. V. Viswanathan, S. Baskaran, V. Sprenkle, X. Li, Y. Shao and B. Schwenzer, *Adv. Funct. Mater.*, 2013, **23**, 929.
- 4 M. S. Islam and C. A. Fisher, *Chem. Soc. Rev.*, 2014, **43**, 185.
- 5 V. V. Politaev, A. A. Petrenko, V. B. Nalbandyan, B. S. Medvedev and E. S. Shvetsova, *J. Solid State Chem.*, 2007, **180**, 1045.
- 6 R. J. Gummow and Y. He, *J. Power Sources*, 2014, **253**, 315.
- 7 A. R. Armstrong, C. Sirisopanaporn, P. Adamson, J. Billaud, R. Dominko, C. Masquelier and P. G. Bruce, *Z. Anorg. Allg. Chem.*, 2014, **640**, 1043.
- 8 C. Masquelier and L. Croguennec, *Chem. Soc. Rev.*, 2013, **113**, 6552.
- 9 G. Mail, C. Sirisopanaporn, C. Masquelier, D. Hanzel and R. Dominko, *Chem. Mater.*, 2011, **23**, 2735.
- 10 A. Boulineau, C. Sirisopanaporn, R. Dominko, A. R. Armstrong, P. G. Bruce and C. Masquelier, *Dalton Trans.*, 2010, 6310.
- 11 R. Dominko, M. Bele, M. Gaberscek, A. Meden, M. Remskar and J. Jamnik, *Electrochem. Commun.*, 2006, **8**, 217.
- 12 M. S. Islam, R. Dominko, C. Masquelier, C. Sirisopanaporn, A. R. Armstrong and P. G. Bruce, *J. Mater. Chem.*, 2011, **21**, 9811.
- 13 M. E. Arroyo-de Dompablo, M. Armand, J. M. Tarascon and U. Amador, *Electrochem. Commun.*, 2006, **8**, 1292.
- 14 M. E. Arroyo-de Dompablo, R. Dominko, J. M. Gallardo-Amores, L. Dupont, G. Mali, H. Ehrenberg, J. Jamnik and E. Morán, *Chem. Mater.*, 2008, **20**, 574.
- 15 S. Nishimura, S. Hayase, R. Kanno, M. Yashima, N. Nakayama and A. Yamada, *J. Am. Chem. Soc.*, 2008, **130**, 13212.
- 16 C. Sirisopanaporn, R. Dominko, C. Masquelier, A. R. Armstrong, G. Maliad and P. G. Bruce, *J. Mater. Chem.*, 2011, **21**, 17823.
- 17 R. J. Gummow, N. Sharma, V. K. Peterson and Y. He, *J. Solid State Chem.*, 2012, **188**, 32.
- 18 H. Duncan, A. Kondamreddy, P. H. J. Mercier, Y. Le Page, Y. Abu-Lebdeh, M. Couillard, P. S. Whifield and I. J. Davidson, *Chem. Mater.*, 2011, **23**, 5446.
- 19 C. Eames, A. R. Armstrong, P. G. Bruce and M. S. Islam, *Chem. Mater.*, 2012, **24**, 2155.
- 20 A. Saracibar, A. Van der Ven and M. E. Arroyo-de Dompablo, *Chem. Mater.*, 2012, **24**, 495.
- 21 C. A. J. Fisher, N. Kuganathan and M. S. Islam, *J. Mater. Chem. A*, 2013, **1**, 4207.
- 22 P. F. Peterson, M. Gutmann, T. Proffen and S. J. L. Billinge, *J. Appl. Crystallogr.*, 2000, **33**, 1192.
- 23 C. L. Farrow, P. Juhas, J. W. Liu, D. Bryndin, E. S. Bozin, J. Bloch, T. Proffen and S. J. L. Billinge, *J. Phys.: Condens. Matter*, 2007, **19**, 335219.

