

Cite this: *RSC Adv.*, 2019, 9, 34430

Diphenyl polysulfides: cathodes with excellent lithiation performance and high specific energy for LSBs†

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Reversible lithium–sulfur batteries (LSBs) are considered one of the most promising next-generation energy storage systems. However, the shuttling effect of lithium polysulfide significantly weakens the electrochemical properties and the cycle life, hindering its practical application. Organo-sulfides are unique materials with low cost, profuse content and high capacity. Here, *via* quantum chemical calculations, we introduce a class of diphenyl polysulfides, PhS_nPh ($2 \leq n \leq 15$), which are all structurally stable, confirmed by calculation of their Gibbs free energies. The theoretical specific energy of PhS_{15}Ph is high, up to 2632 W h kg^{-1} , exceeding that of S_8 . By calculating the bond dissociation energy of S–S in PhS_nPh molecules, we analyze the breaking processes of the S–S bonds in each step of lithiation. The microscopic mechanism of the fast reaction kinetics of PhS_nPh cathodes is explored. It is phenyl that prevents the formation of soluble long-chain polysulfide molecules (Li_2S_4 , Li_2S_6 , Li_2S_8) in the lithiation process, efficiently weakening the “shuttle effect”.

Received 16th August 2019
Accepted 14th October 2019

DOI: 10.1039/c9ra06402h

rsc.li/rsc-advances

1. Introduction

Compared with traditional lithium ion secondary batteries, rechargeable lithium sulfur batteries (LSBs) are widely considered an ideal choice for electric vehicles due to their high theoretical specific capacity and specific energy.¹ However, many problems and challenges limit the application of LSBs. One of the most important problems is the “shuttle effect”.^{2,3} The capacity attenuation,⁴ low coulomb efficiency⁵ and high self-discharge rate⁶ caused by the “shuttle effect” restrict the practical application of LSBs.⁷ Inhibiting the shuttling of polysulfide is the key to the development of Li–S batteries for excellent electrochemical properties.⁸ Extensive efforts have been made to improve it.^{9,10} Amruth Bhargav *et al.*¹¹ used CNTs as the porous and conductive carbon matrix to efficiently trap soluble polysulfides. Liang X. *et al.*¹² synthesized sulfur/manganese dioxide nanosheet composite which can entrap polysulfides in the cathode. Tae-Gyung Jeong *et al.*¹³ encapsulated sulfur particles with functional polymers, which enhanced the cycling stability by suppressing the dissolution of polysulfide from the sulfur materials and rendered the electrodes less reactive toward liquid electrolyte. Jing Zheng *et al.*¹⁴

proposed a new kind of localized high-concentration electrolyte. It achieved a high coulombic efficiency up to 99.3% and completely suppressing the shuttling effect. Tianyu Lei *et al.*¹⁵ synthesized PAN@APP microfiber separator to bind interaction with polysulfides strongly and improve the safety.

Recently, an excellent cathode material with high cyclic stability and energy efficiency has been studied by Min Wu *et al.*¹⁶ They compounded PhS_3Ph with the PhS_2Ph and S_8 . The cell delivers an initial discharge specific energy of 751 W h kg^{-1} with high energy efficiency. Then, Amruth Bhargav *et al.*¹⁷ introduced a new class of phenyl polysulfides PhS_nPh ($4 \leq x \leq 6$) as liquid cathode materials. Their volume change when reduced is about 37% and PhS_6Ph can provide a specific energy of 1665 W h kg^{-1} . Therefore, increasing the number of S atoms in polysulfide diphenyl could improve the specific capacity of LSBs. Wei Chen *et al.*¹⁸ achieved a cathode with high rate and stable cycling performance by increasing sulfur content.

In this paper, we investigate the stability PhS_nPh ($2 \leq n \leq 15$) molecules by density functional theory calculation. According to the calculation results, the sulfur atoms in the PhS_nPh molecules behave chain-like spiral distribution between two phenyl groups, the structure of PhS_nPh molecules are stable by analyzing the formation of Gibbs free energy. More importantly, the specific energy of PhS_{15}Ph is up to 2632 W h kg^{-1} , very close to the theoretical value of S_8 . Furthermore, we analyze the lithiation process of the PhS_nPh cathodes by calculating the bond dissociation energies of S–S bonds in the PhS_nPh molecules. The calculating results reveal a new kind of inhibition mechanism of “shuttle effect”. In the lithiation process, as an intermediate lithiation product, the long chain PhS_nLi has

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† Electronic supplementary information (ESI) available. See DOI: 10.1039/c9ra06402h



excellent structural stability after phenyls are added as small electrophilic groups. The final lithiation products are PhSLi, Li₂S and Li₂S₂. There's no Li₂S_n ($n = 4, 6, 8$) molecules being introduced in the whole discharge process, which are the major factor for the generating "shuttle effect". In the experiment, the PhS_nPh ($2 \leq n \leq 6$) have been synthesized successfully,^{16,17} the PhS_nPh ($7 \leq n \leq 15$) with more S atoms are confirmed to be thermodynamically stable by the analysis of "alloy-like" diagrams.¹⁹ Considering elemental abundance²⁰ and scalability of synthetic methods, the LSBs with PhS_nPh as cathodes also may surpass lithium-ion battery as a electrochemical storage system.²¹ Therefore, this class of polysulfides could be a kind of promising high-capacity cathode material for LSBs.

2. Computational methods

All the calculations of Gibbs free energies presented in this paper were performed using the Gaussian 16 code. The polarizable continuum model (PCM)²² was used to describe the interaction between PhS_nPh/PhS_nLi and the electrolyte. The fully optimization of geometries, Gibbs free energies, bond dissociation energy (BDE) and vibrational frequency were performed by the B3LYP/6-311+G(d) DFT functional and basis set, which were widely used to calculate the organo-sulfides.²³ The electronic excitation energies were calculated at the TD-B3LYP/6-311+G(d) level of theory. Gaussian band-shapes with a bandwidth of 0.2 eV (two thousand wavenumbers) were used to simulate the UV-vis spectra.

In order to confirm the possibility of adding more sulfur atoms between two phenyls, here we make an "alloy-like" diagram¹⁹ to analyze the possibility. The Gibbs free energy difference (ΔG) is used to analyze the relative stability and reaction feasibility of PhS_nPh, the formation energies are defined as:

$$G_{\text{form}} = \left[G_{\text{PhS}_n\text{Ph}} - \frac{1}{8}(n-2)G_{\text{S}_8} - G_{\text{PhS}_2\text{Ph}} \right] / (n-1) \quad (1)$$

As shown in Fig. 1, when $n = 0, 1$, they stand for the reactant PhS₂Ph and S₈, respectively, the G_{form} equals zero. It can be seen that all of G_{form} are negative. As the number of S increases, G_{form} decreases with the number of S atoms increasing, indicating that more sulfur atom could be added between two phenyls. The advantage of "alloy-like" diagram is to see whether a particular compound is thermodynamically stable. If one compound is above the convex hull, it will be unstable, and will decompose into the two nearby compounds. Judging by the results shown in Fig. 1, all the diphenyl polysulfides are thermodynamically stable.

3. Results and discussion

3.1 PhS_nPh molecules

Firstly, the structures of PhS_nPh molecules are optimized without considering the solvent effect. After completely optimized, phenyl groups are at both ends, and sulfur atoms are chain-like between the two phenyl groups. The bond lengths, bond angles of S-S bonds and S-Ph bonds, and dihedral angles of phenyls are listed in Table 1. The S-S bond lengths are in the interval of 2.04–2.09 Å and the interval of bond angles of S-S bonds are 107.48–110.26°, which in good agreement with the literature of Ralf Steude's review about the chain-like structures of RS_nR molecules.²⁴ In Fig. 1, we show the optimized structural configures of PhS_nPh molecules, the S-S bonds between two benzene ring screw around in the range of 320–330°, close to 360°, so the two phenyl groups nearly symmetric at $n = 3, 6, 9, 12$. This feature can also be verified from the dihedral angles. The dihedral angles of PhS₃Ph, PhS₆Ph, PhS₉Ph, PhS₁₂Ph are smaller, which are 1.38°, 40.32°, 26.74° and 16.77°, respectively.

To compare some useful information with *in situ* and operando experiments, we simulated ultraviolet-visible (UV-vis) spectroscopy. As shown in Fig. 2(a), there is only one peak among the ultraviolet range (200–700 nm) for different phenyl polysulfides; the peak wavelength (λ , in nm) increase

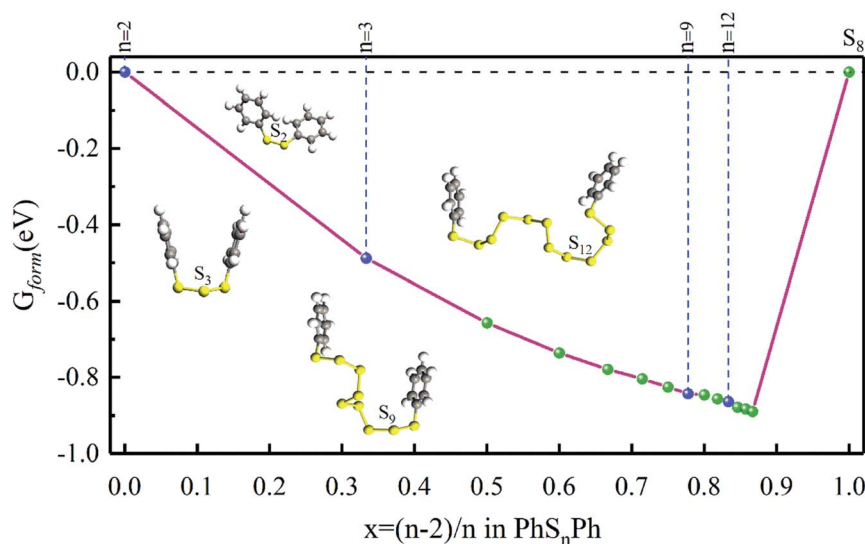


Fig. 1 The G_{form} for PhS_nPh ($1 \leq n \leq 15$), the inset show the optimized stable structures of PhSPh, PhS₂Ph, PhS₃Ph, PhS₉Ph and PhS₁₂Ph, respectively.



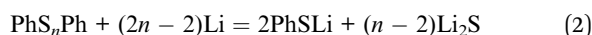
Table 1 The bond lengths, bond angles and dihedral angles of PhS_nPh (1 ≤ *n* ≤ 12)

	Bond length <i>d</i> _{SS} (Å)	Bond angle <i>α</i> _{SSS} (deg)	Dihedral angle <i>γ</i> _{Ph} (deg)	Bond length <i>d</i> _{SPh} (Å)	Bond angle <i>α</i> _{SSPh} (deg)
PhSPh	—	—	3.36	1.78	140.37
PhS ₂ Ph	2.04	—	47.79	1.79	107.02–107.04
PhS ₃ Ph	2.07–2.08	109.75	1.38	1.79	105.56–106.25
PhS ₄ Ph	2.08	109.12–109.29	77.14	1.78	104.64–105.04
PhS ₅ Ph	2.07–2.08	108.62–108.64	128.76	1.78	104.55–105.31
PhS ₆ Ph	2.07–2.08	108.58–108.96	40.32	1.78	104.60–105.02
PhS ₇ Ph	2.07–2.08	108.31–109.58	128.13	1.79	105.07–105.74
PhS ₈ Ph	2.07–2.09	107.88–109.63	146.72	1.78	104.76–105.70
PhS ₉ Ph	2.07–2.08	108.01–109.19	26.74	1.78	104.69–105.12
PhS ₁₀ Ph	2.07–2.09	107.78–110.26	136.86	1.78	105.08–107.20
PhS ₁₁ Ph	2.07–2.09	107.51–110.15	80.13	1.78	104.75–105.13
PhS ₁₂ Ph	2.07–2.09	107.48–110.07	16.77	1.78	104.67–105.05

with the number of S atom (*n*) and convergence at the PhS₆Ph. Only considering the PhS₂Ph, PhS₄Ph, PhS₅Ph and PhS₆Ph, the curve of peak λ behave almost linear strand, which is in good agree with the experiment dates.¹⁷ When *n* (number of sulfur atoms) is an integer multiple of 3 (*n* = 3, 6, 9 and 12), the peak λ behaves a bigger bathochromic shift which are shown with red dots in Fig. 2(b). The peak wavelengths are derived from the out-of-plane deformation of the phenyl rings due to the polysulfide linkages.²⁵ According to the structure configures shown in inset Fig. 1, the sulfur atoms between two phenyls are chain-like and present a spiral, three sulfur atoms form a unit; when there are integer units in the linkages (*n* = 3, 6, 9 and 12), the dihedral angles are smaller, the structure configures of PhS₃Ph, PhS₆Ph, PhS₉Ph, PhS₁₂Ph show more symmetric and behave more stable. The characteristics of the above-mentioned UV-vis spectroscopy of PhS_nPh could be used to identify the synthesized polysulfides in experiments.

3.2 Average voltage and specific energy of PhS_nPh cathodes

In order to calculate the average voltages of PhS_nPh cathodes, the implicit solvent model (PCM) was used to describe the interaction between PhS_nPh and the electrolyte. Both in experiment²⁶ and calculation,²⁷ 1,2-dimethoxyethane (DME) and 1,3-dioxolane (DOL) are the electrolytes widely used in lithium-sulfur batteries at present. With DME/DOL (1 : 1, v/v), the theoretical voltage of polysulfides is quite close to the experimental voltage of 2.0–2.6 V.^{28,29} Therefore, in the following research, all calculation consider solvation effect of DME/DOL (1 : 1, v/v) to make our results more reasonable and close to reality. We calculate average voltages according to the average energy changes of full lithiation processes.³⁰ The process of full lithiation can be expressed as:



the average free energy change per step of lithiation is:

$$G_{\text{ave}} = [2G_{\text{PhSLi}} + (n - 2)G_{\text{Li}_2\text{S}} - G_{\text{PhS}_n\text{Ph}} - (2n - 2)G_{\text{Li}}]/(2n - 2) \quad (3)$$

Note that, not existing originally in the anode, the Gibbs free energy here for Li is electrolyte-free. The average voltages of PhS_nPh cathodes are shown in Fig. 3, as well as the discharge specific capacity and specific energy.

From Fig. 3, the average voltages first reduce quickly as *n* increases in PhS_nPh, and then approach to 2.227 V when *n* ≥ 9. Next, we analyze the discharge specific capacities and specific energies. With the increasing of S atoms, they keep rising. The specific energy of PhS₁₅Ph, 2632 W h kg^{−1}, reaches the theoretical value of S₈, 2600 W h kg^{−1}.³¹ And its specific capacity is

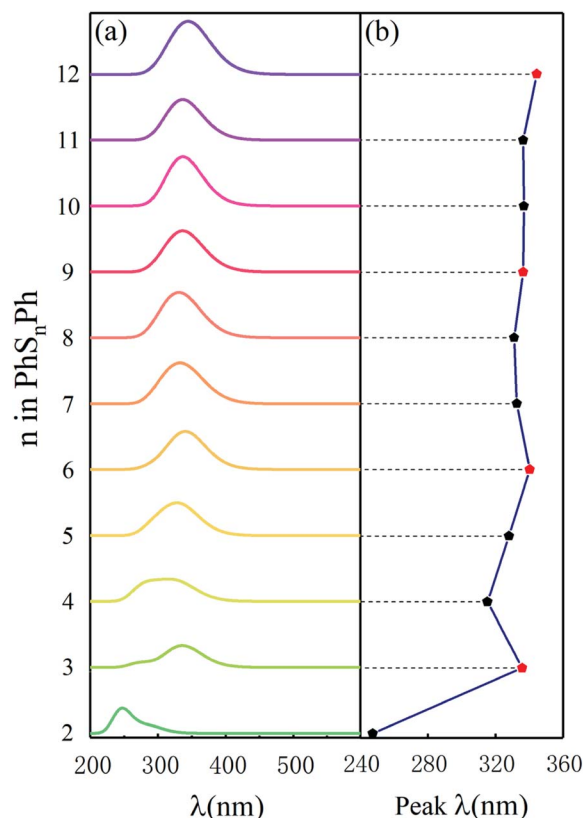


Fig. 2 (a) Simulated UV-vis spectra and (b) the peak λ in UV-vis spectra of the PhS_nPh (2 ≤ *n* ≤ 12) molecules.



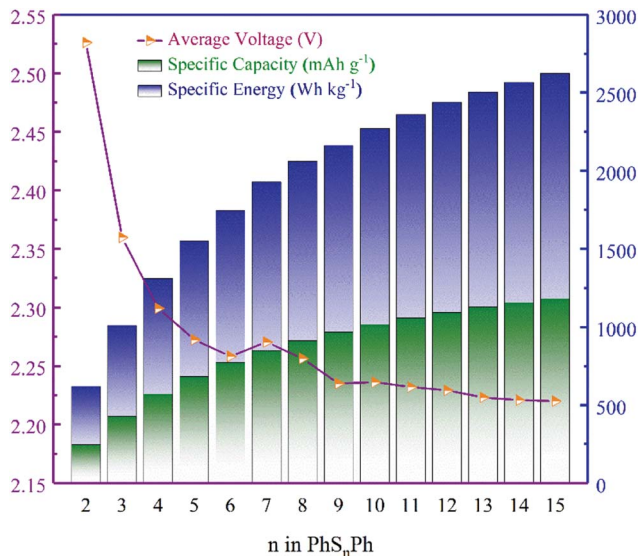


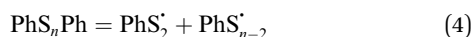
Fig. 3 The average voltages, specific capacities and specific energies of the PhS_nPh cathode.

up to 1182 mA h g^{-1} , nearly five times that of state-of-the-art cathode materials used in Li-ion batteries.³² The comparison of theoretical specific capacities and energies clearly indicate the great promise PhS_nPh hold in surpassing the current cathode materials used in Li-ion batteries.

3.3 The discharge lithiation process of PhS_nPh cathodes

Here we propose the bond dissociation energy (BDE) as another method to analysis the lithiation. BDE, also called bond dissociation enthalpy, is the most effective quantitative description of chemical bonds.³³ It is defined as the change in reaction enthalpy of the breaking process of chemical bonds, that is, BDEs are investigated by the density functional theory.³⁴ In this way, DFT can predict excellent results for these energies.³⁵ The sum of electronic and thermal enthalpies of PhS_nPh , $\text{PhS}_x^\bullet$ and $\text{PhS}_{n-x}^\bullet$ are calculated to determine the dissociation energies of the bonds. BDEs are all positive because the bonds need to absorb energies to dissociate. In other words, it's easier for bonds to break where the dissociation energy is lower. M. J. Bausch *et al.* has estimated S-S BDE of phenyl disulfide. Literature data indicates that the gas-phase BDE for the S-S bonds in PhSSPh is about 55 kcal mol^{-1} , 2.31 eV .³⁶ With the method above, S-S BDE of PhS_2Ph in DME/DOL (1 : 1, v/v) is 2.46 eV . It is a little bigger than that of gas-phase because of different states. But they are very consistent and enough to demonstrate that our calculation is reliable.

According to Fig. 4, the first step of lithiation may react at the second S-S bond or the penultimate one, which can be described equally as:



What is more, it could be that both of the two bonds break at the same time:

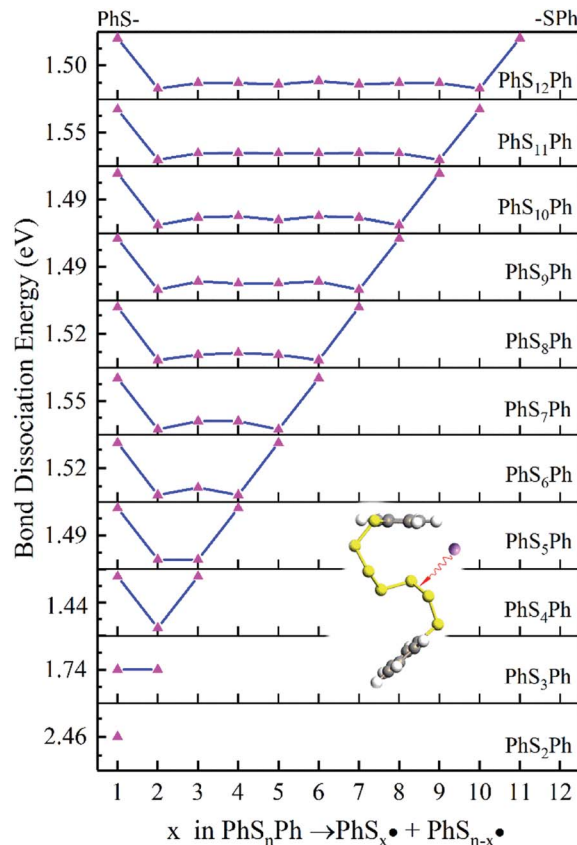
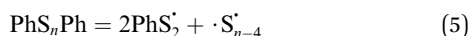
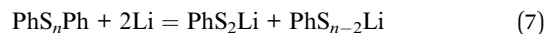


Fig. 4 The BDE of S-S bonds in the PhS_nPh . The inset shows the breaking position in PhS_6Ph .

In judging the strength of a bond in a chemical reaction, we should also consider the stability of free radicals after dissociation. Since the free radicals would directly generate PhS_nLi with Li^+ , the stability of the $\text{PhS}_x^\bullet$ can be measured by the Gibbs formation energies of PhS_nLi . We also use an “alloy-like” diagram to analyze the optimized molecules. We get the formation energies based on this equation:

$$G_{\text{form}} = [G_{\text{PhS}_n\text{Li}} - (n-1)G_{\text{S}_8/8} - G_{\text{PhSLi}}]/n \quad (6)$$

As it is shown in Fig. 5, long-chain PhS_nLi ($n \geq 3$) is more stable than PhS_2Li with lower Gibbs formation energy. So in the first step of lithiation, only one bond would dissociate to form PhS_2Li and $\text{PhS}_{n-2}\text{Li}$. Concern is that there is no Li_2S_n in this reaction:



We use the same ways to analysis the subsequent lithiation of PhS_nLi , shown in Fig. 6. The lower position of dissociation energy of S-S bond increases with the increase of n . But considering better stability of long-chain PhS_nLi ($n \geq 3$), the lithiation of PhS_nLi would react at the penultimate S-S bond to generate $\text{PhS}_{n-2}\text{Li}$ and Li_2S_2 ($n \geq 3$) or Li_2S ($n = 2$). The latter lithiations will follow this pattern until they are completely reduced into PhSLi and Li_2S .

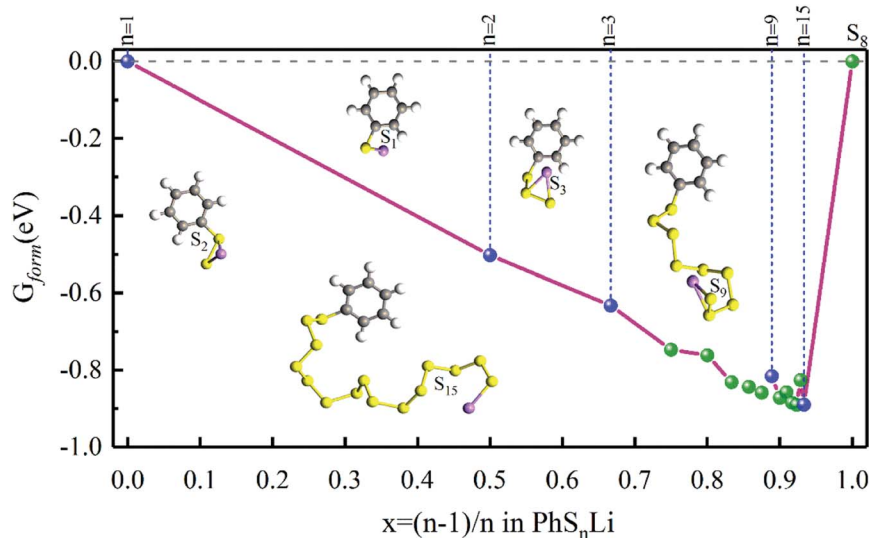


Fig. 5 The G_{form} for PhS_nLi . The insets show the optimized stable structures of PhS_1Li , PhS_2Li , PhS_3Li , PhS_9Li and PhS_{15}Li , respectively.

In conclusion, at the beginning of the discharge, PhS_nPh is reduced to PhS_2Li and $\text{PhS}_{n-2}\text{Li}$. With the depth of discharge increasing, $\text{PhS}_{n-4}\text{Ph}$, $\text{PhS}_{n-6}\text{Ph}$, $\text{PhS}_{n-8}\text{Ph}$ and other intermediate products (PhS_xLi) are generated in sequence. Meanwhile, Li_2S_2 is formed. Finally, PhS_2Li is further reduced to PhSLi and Li_2S . Li_2S_2 is deoxidized into Li_2S . This chemical change can be expressed by four steps:

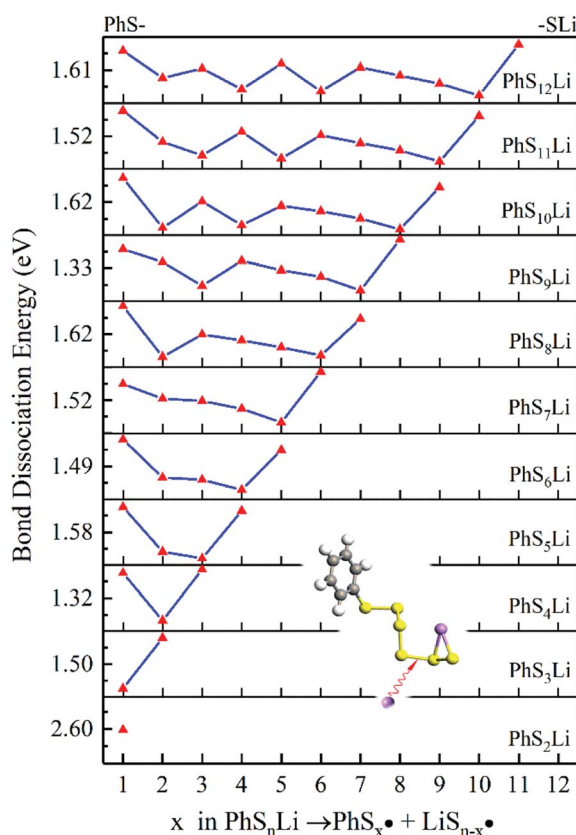
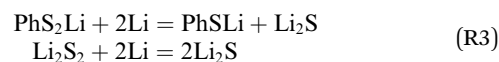
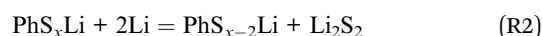
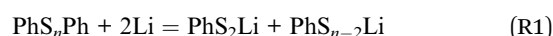


Fig. 6 The BDE of S-S bonds of PhS_nLi . The inset shows the breaking position of PhS_6Li .

In this discharge sequence, the reaction energies (G_{rec}) can be calculated with the energy difference between the reactants and the products. Based on reaction energies, the corresponding voltages *versus* Li/Li^+ are listed in Table S1.† Fig. 7(a) shows calculated reaction sequences of PhS_5Ph . PhS_5Ph may break the S-S bond and form PhS_3Li and PhS_2Li when attracted by Li^+ and e^- firstly. During the step of discharge, the reaction energy is 5.05 eV and the voltage is 2.53 V. Then PhS_3Li will take another Li^+ and e^- to generate PhS_2Li with 2.52 V. The final step is that PhS_2Li and Li_2S_2 are reduced into PhSLi and Li_2S , respectively. The two reactions can be regarded as the final step of the whole lithiation and the average voltage is 2.04 V. The voltage profiles of the entire discharge process are mostly consistent with the experiments.¹⁷

To reveal the structure change of PhS_nPh when attacked by a Li^+ and a e^- , we take PhS_5Ph for example to draw the difference map of electron density. In Fig. 7(b), red solid lines and blue dashed lines correspond to the regions having increased electron density and decreased electron density during the process, respectively. Obviously, the $\text{S}_2\text{-S}_3$ bond is weakened sharply and the Li^+ is bonded to S_3 . Furthermore, the $\text{S}_2\text{-S}_3$ bond distance increase from 2.08 Å to 3.19 Å. PhS_5Ph molecule splits up to form PhS_3Li and $\text{PhS}_2^\bullet$ radical, the latter of which will further obtain a Li^+ and a e^- to form PhS_2Li . The result is the same with the first lithiation step derived from BDE. For this reason, it also proves that our method of analyzing the lithiation steps by BDE is reliable.

The discharge order of PhS_{10}Ph is shown in Table 2. In the first step, PhS_{10}Ph is reduced to PhS_2Li and PhS_8Li at 2.68 V. As the discharge depth increases, PhS_8Li is reduced to PhS_6Li , PhS_4Li and PhS_2Li in sequence with voltages of 2.52 V, 2.46 V



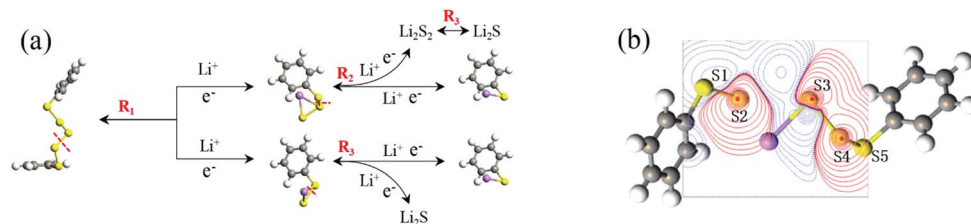


Fig. 7 (a) Schematic illustration of the discharge process of the PhS_5Ph cathode; (b) relaxed structure of $\text{PhS}_5\text{Ph} + \text{Li}$, the difference map of electron density between S and Li are also shown.

Table 2 The discharge reactions, reaction energies and corresponding voltages of PhS_5Ph and PhS_{10}Ph

Label	Category	Reactions	G_{rec} (eV)	Voltage (V)	Feasibility
R ₁	PhS_5Ph	$\text{PhS}_5\text{Ph} + 2\text{Li} = \text{PhS}_2\text{Li} + \text{PhS}_3\text{Li}$	−5.05	2.53	✓
R ₂		$\text{PhS}_3\text{Li} + 2\text{Li} = \text{PhSLi} + \text{Li}_2\text{S}_2$	−5.03	2.52	✓
R ₃		$\text{PhS}_2\text{Li} + 2\text{Li} = \text{PhSLi} + \text{Li}_2\text{S}$	−4.36	2.18	✓
R ₃		$\text{Li}_2\text{S}_2 + 2\text{Li} = 2\text{Li}_2\text{S}$	−3.80	1.90	✓
R ₁	PhS_{10}Ph	$\text{PhS}_{10}\text{Ph} + 2\text{Li} = \text{PhS}_2\text{Li} + \text{PhS}_8\text{Li}$	−5.36	2.68	✓
R ₂		$\text{PhS}_8\text{Li} + 2\text{Li} = \text{PhS}_6\text{Li} + \text{Li}_2\text{S}_2$	−5.04	2.52	✓
R ₂		$\text{PhS}_6\text{Li} + 2\text{Li} = \text{PhS}_4\text{Li} + \text{Li}_2\text{S}_2$	−4.92	2.46	✓
R ₂		$\text{PhS}_4\text{Li} + 2\text{Li} = \text{PhS}_2\text{Li} + \text{Li}_2\text{S}_2$	−4.94	2.47	✓
R ₃		$\text{PhS}_2\text{Li} + 2\text{Li} = \text{PhSLi} + \text{Li}_2\text{S}$	−4.36	2.18	✓
R ₃		$\text{Li}_2\text{S}_2 + 2\text{Li} = 2\text{Li}_2\text{S}$	−3.80	1.90	✓

and 2.47 V, respectively. Finally, PhS_2Li and Li_2S_2 are reduced into PhSLi and Li_2S at an average voltage of 2.04 V. The voltage prediction will be of great value in the research of PhS_nPh as cathodes for Li–S batteries.

In the whole discharge process, there is no Li_2S_4 , Li_2S_6 , Li_2S_8 . To compare the dissolution of PhS_nLi and Li_2S_n ($n = 4, 6, 8$) in the DOL/DME solvent, the dissolving free energies³⁷ of them with the solvent effect of DOL/DME (1 : 1, v/v) are calculated. As shown in Fig. S1,† the dissolving free energies of PhS_nLi ($1 \leq n \leq 13$) are almost above those of Li_2S_n ($n = 4, 6, 8$) which result in shuttle effect. Therefore, the solubilities of PhS_nLi ($1 \leq n \leq 13$) are lower and the shuttle effect caused by PhS_nLi is much weaker than that by Li_2S_n ($n = 4, 6, 8$). We take it that the excellent stability of long-chain PhS_nLi and the dissociation positions of S–S bonds jointly contribute to such a result. Here, it can be considered that phenyl plays a certain role in fixing soluble polysulfide, preventing polysulfide from being dissolved into the electrolyte, which can greatly slow down the “shuttle effect”. This is beneficial to improve coulomb efficiency and capacity attenuation, and achieve better cycle stability and energy efficiency in practical applications. Until now, there are some ways to eliminate the “shuttle effect” particularly: improve the electrolyte performance,³⁸ recombine cathodes with organic polymers³⁹ or porous materials⁴⁰ and so on. In comparison to these ways, the synthesis of PhS_{15}Ph is more facile and available. So this class of cathode materials is of great research significance and potential for LSBs.

4. Summary

In conclusion, using the *ab initio* density functional theory calculation, we have performed thorough theoretical studies

for diphenyl polysulfides as cathode materials for LSBs. One goal is to increase the sulfur proportion in PhS_nPh compounds, more sulfur proportion means high specific capacity; the other goal is to overcome the “shuttle effect” challenge for the current LSBs, namely the dissolution of lithium polysulfide. Though the Gibbs free energy calculation and “alloy-like” diagram analysis, we found the PhS_nPh molecules are thermodynamically stable, and the PhS_{15}Ph behave almost the same energy density as S_8 , 2632 W h kg^{−1}. By calculation the BDE of the PhS_nPh cathode, we found that only short-chain polysulfides generate in the discharge process, PhS_nPh prevents the formation of soluble long-chain intermediates that plague traditional sulfur cathodes. We hope that this work leads to further studies in experiment, which would make PhS_nPh a promising candidate for low-cost, eco-friendly, and intrinsically safe cathode materials for applications in rechargeable lithium batteries.

There are other aspects which we haven't discussed until now, most importantly, the volumetric capacity. One cannot get high volumetric capacity by using only one PhS_nPh molecule, a possible solution might include the use of PhS_nPh molecule to establish unit cell of 3D frameworks,⁴¹ but poor conductivity is a very big defect. Recently, Se is introduced into S cathodes by forming Se–S bonds to modify the electronic and ionic conductivity and ultimately enhance cathode utilization in LSBs,^{42,43} this is also the topic we need to study next.

Conflicts of interest

There are no conflicts to declare.



Acknowledgements

JBW was supported by the National Natural Science Foundation of China (No. 11047164), the Shanghai College Foundation for Excellent Young Teachers of China (No. gjd10023) and the Academic Program of Shanghai Municipal Education Commission (No. 11XK11 and 2011X34).

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