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A novel self-assembled supramolecular sensor based on thiophene-functionalized imidazophenazine for dual-channel detection of Ag⁺ in an aqueous solution†

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Herein, a novel self-assembled supramolecular sensor (**S1**) based on thiophene-functionalized imidazophenazine for Ag⁺ was designed and synthesized. It showed dual-channel detection properties for Ag⁺ based on the competitive mechanism of supramolecular self-assembly with high sensitivity and selectivity even in the presence of other metal ions. Its detection limit for Ag⁺ is 8.18×10^{-9} M. Upon the addition of Ag⁺, the solution changes from yellow to light purple and the fluorescence is quenched. Furthermore, the sensing mechanism between Ag⁺ and **S1** is investigated via IR and ¹H NMR spectroscopy, mass spectrometry, and density functional theory calculations.

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Introduction

With the development of modern industry, increasing industrial emissions containing metal ions have caused serious water pollution, which affects food and agriculture and further threatens human health and the environment.¹ Among the important metal ions, silver has attracted extensive attention due to its wide applications in the fields of electricity, photography, imaging, and pharmacy.² In fact, trace amounts of silver ions can control bacteria breeding or rehabilitate and reconstruct essential human tissue.³ However, as is well known, excess silver ions can inactivate sulfhydryl enzymes; on the other hand, silver ions can combine with amine, imidazole, and carboxyl groups of various metabolites, such as high molecular weight protein and metallothionein, in the tissue of the cytosol fraction, resulting in harmful effects to humans.⁴ Thus, the maximum bearable concentration of Ag(I) issued by the World Health Organization (WHO) is 0.5 μM in drinking water.⁵

Considering the importance of Ag(I) in industrial production and in daily life, the development of techniques for sensing and monitoring Ag(I) is in great demand. There are various available methods to monitor Ag⁺, such as electrothermal atomic absorption spectrometry (ETAAS), voltammetry, inductively coupled plasma atomic emission spectrometry (ICP-AES), inductively coupled plasma mass spectroscopy (ICP-MS), and

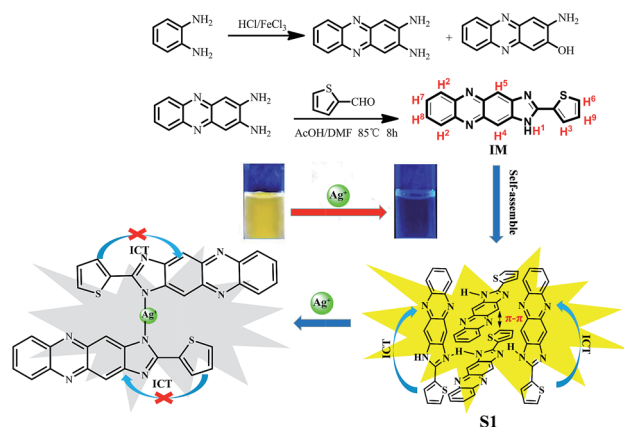
potentiometry.⁶ However, most of them are expensive, time-consuming, or unsuitable for on-site and real-time monitoring. Therefore, the development of fast, efficient, sensitive, and selective methods for the detection of Ag⁺ is necessary for both the environment and human health.⁷ In comparison, fluorescence techniques exhibit attractive advantages, such as high selectivity and sensitivity, real-time detection, practical-simple sample preparation, and great potential in high throughput detection of Ag⁺ on-site. However, some of these probes suffer from serious drawbacks such as complicated synthesis procedures and poor solubility, sensitivity, and selectivity in aqueous media.⁸ Due to the efforts of researchers, numerous Ag⁺ sensors with good properties have been reported. However, most of them exhibit ultraviolet or fluorescent single channel recognition or are less selective.⁹ Moreover, it is worthwhile to note that sensors for the detection of Ag⁺ in aqueous systems are relatively scarce although Ag⁺ usually exists in water or sewerage. Consequently, the design and synthesis of colorimetric and fluorescent chemosensors with high sensitivity and selectivity for the quantification of Ag⁺ in aqueous media via a simple reaction is still an intriguing area of research.¹⁰

In view of this and based on our previous research in host-guest chemistry,¹¹ we designed and synthesized a sensor molecule **IM** (Scheme 1) based on thiophene-functionalized imidazophenazine via a simple reaction; this sensor could self-assemble into a supramolecular system and form a supramolecular chemosensor **S1**. This sensor combines the phenazine fluorescent signaling subunit, which has attracted extensive attention due to its excellent photophysical properties. In addition, among the different fluorogens, it is very sensitive to

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Scheme 1 Synthesis of the compound IM and the proposed possible mechanism of **S1** bonding with Ag^+ .

conformational changes.¹² Consequently, phenazines are ideal platforms for the development of cation, anion, and neutral molecule recognition platforms. Moreover, there are only a few reports on ion recognition in host–guest chemistry. Compared to some Ag^+ sensors reported in previous studies, the Ag^+ sensor reported herein is a self-assembled supramolecular sensor (**S1**). The Ag^+ sensing mechanism is based on a competition between self-assembly of the sensor molecules and Ag^+ coordination with the sensor molecules. This novel competitive mechanism provides the **S1** sensor with better Ag^+ detection sensitivity. The detection limit of **S1** for Ag^+ is 8.18×10^{-9} M, which is lower than that reported in some previous studies.⁸ Therefore, **S1** is utilized as a supramolecular self-assembled sensor to study the recognition performance of Ag^+ .

Results and discussion

To investigate the Ag^+ recognition abilities of the **S1** sensor in an aqueous solution, we performed a series of host–guest recognition experiments. The recognition profiles of the chemosensor **S1** towards various metal ions, including Fe^{3+} , Hg^{2+} , Ag^+ , Ca^{2+} , Cu^{2+} , Co^{2+} , Ni^{2+} , Cd^{2+} , Pb^{2+} , Zn^{2+} , Cr^{3+} , and Mg^{2+} , were primarily investigated using UV-vis spectroscopy and fluorescence spectroscopy in a DMSO/water (v:v = 1 : 1, buffered with HEPES pH = 7.2) solution.

As shown in Fig. 1, the sensor immediately responded with obvious color changes (from yellow to light purple) when Ag^+ was added to the **S1** solution at room temperature. In the corresponding UV-vis spectra, the strength of the absorption peak at 413 nm decreased, and the peak red-shifted to 376 nm (Fig. S1†). However, yellow fluorescence with one emission band centered at 547 nm appeared when the **S1** sensor solution was excited at 400 nm. The fluorescence intensity of the peak at 547 nm sharply decreased with the addition of the Ag^+ solution. No significant UV-vis or fluorescence spectra changes were observed when solutions of other cations (such as Fe^{3+} , Hg^{2+} , Ca^{2+} , Cu^{2+} , Co^{2+} , Ni^{2+} , Cd^{2+} , Pb^{2+} , Zn^{2+} , Cr^{3+} , and Mg^{2+}) were added to the **S1** sensor solution.

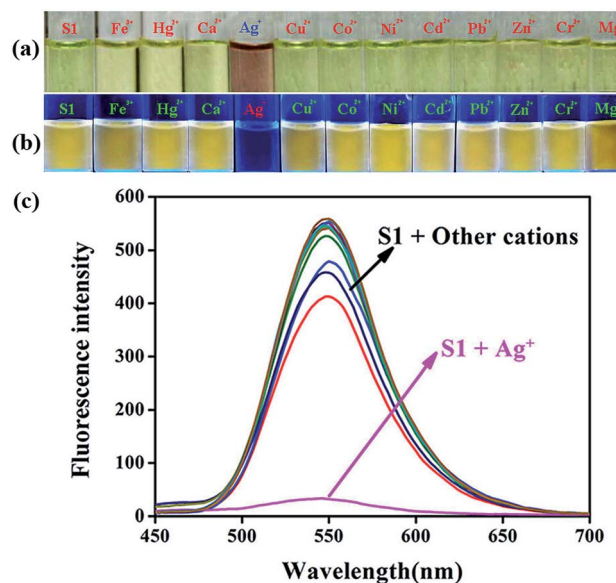


Fig. 1 An image of **S1** upon the addition of 10 equiv. of various cations in DMSO/ H_2O (v:v = 1 : 1, buffered with HEPES pH = 7.2), which was taken under (a) visible light and (b) UV-lamp (365 nm) at room temperature. (c) Fluorescence spectral responses in DMSO/ H_2O (v:v = 1 : 1, buffered with HEPES pH = 7.2) of **S1** (2.0×10^{-5} M) upon the addition of different cations (10 equiv.).

Subsequently, we further investigated the supramolecular sensor **S1** recognition for Pd^{2+} , as shown in Fig. S2 and S3;† however, **S1** could not sense Pd^{2+} . These results confirmed that the self-assembly supramolecular sensor **S1** could sense Ag^+ with high sensitivity and selectivity.

The specific selectivity of a sensor towards an analyte in the presence of other competitive species is important. To explore the utility of the **S1** sensor as an ion-selective chemosensor for Ag^+ , competitive experiments were carried out in the presence of 10 equiv. of Ag^+ and 10 equiv. of Fe^{3+} , Hg^{2+} , Ca^{2+} , Cu^{2+} , Co^{2+} , Ni^{2+} , Cd^{2+} , Pb^{2+} , Zn^{2+} , Cr^{3+} , and Mg^{2+} metal ions in DMSO/water (v:v = 1 : 1, buffered with HEPES pH = 7.2) binary solutions of the **S1** sensor (Fig. 2 and S4†). From the bar diagram, it was observed that none of these ions induced any significant change in the fluorescence and UV-vis absorbance spectrum of the sensor. Therefore, it is clear that the interference of other ions is negligible during the detection of Ag^+ . These results further suggested that **S1** could be used as a sensor for Ag^+ over a wide range of cations. Thus, **S1** could selectively and instantly (dual-channel) detect Ag^+ in DMSO/water (v:v = 1 : 1, buffered with HEPES pH = 7.2).

To evaluate the Ag^+ -responsive nature of **S1**, fluorescence and UV-vis absorbance titrations with Ag^+ at varying concentrations were conducted. As shown in Fig. 3, upon the addition of Ag^+ to the **S1** sensor, the emission peak at 547 nm gradually diminished, and the fluorescence of **S1** was almost quenched by 18.0 equiv. of Ag^+ ions. In the UV-vis spectra (Fig. 4), the initial absorption peak at 413 nm gradually decreased, and simultaneously, a broad shoulder peak appeared at 530 nm and an isosbestic point at 495 nm was clearly observed. This should be attributed to the damage of intramolecular hydrogen-bonding



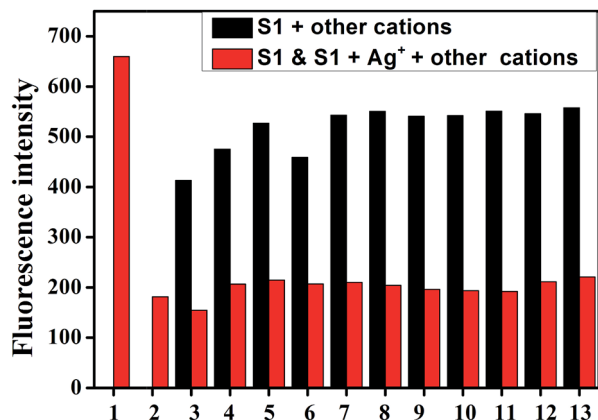


Fig. 2 Fluorescent emission ($\lambda_{\text{ex}} = 400$ nm) spectra of **S1** (2×10^{-5} M) and Ag^+ (10 equiv.) in the presence of various competing cations in DMSO/ H_2O (v:v = 1 : 1, buffered with HEPES pH = 7.2) solution. From 1 to 13: **S1**, **S1** + Ag^+ , **S1** + Ag^+ + Fe^{3+} , **S1** + Ag^+ + Ca^{2+} , **S1** + Ag^+ + Cu^{2+} , **S1** + Ag^+ + Co^{2+} , **S1** + Ag^+ + Ni^{2+} , **S1** + Ag^+ + Cd^{2+} , **S1** + Ag^+ + Pb^{2+} , **S1** + Ag^+ + Zn^{2+} , **S1** + Ag^+ + Hg^{2+} , **S1** + Ag^+ + Cr^{3+} , and **S1** + Ag^+ + Mg^{2+} , respectively.



Fig. 3 Fluorescence spectra of **S1** (20 μM) in the presence of different concentrations of Ag^+ (0–18 equiv.) in DMSO/ H_2O (v:v = 1 : 1, buffered with HEPES pH = 7.2) solution. Inset: A plot of the fluorescence intensity depending on the concentration of Ag^+ in the range from 0 to 18 equiv.

interactions in the sensor and confirms the coordination of the sensor and Ag^+ . Based on the equation $\text{LOD} = K \times \delta/S$,¹³ the detection limits obtained from the UV-vis and the fluorescence spectra were 8.75×10^{-7} M and 8.18×10^{-9} M for silver ions, respectively.¹⁴ This data indicates that the sensor can detect Ag^+ at very low concentrations in the environment with high selectivity and sensitivity. This result also shows that **S1** has higher sensitivity for silver ions as compared to other reported Ag^+ sensors (Table S1†). Therefore, **S1** can be used as a potential colorimetric and fluorescent sensor for Ag^+ recognition.

Since the pH value of the environment around a chemosensor usually influences its detection toward target metal ions due to the protonation or deprotonation reaction of the sensor, the influence of pH value on the absorbance and emission bands of **S1** in a DMSO/water (1 : 1, v:v, buffered with HEPES pH

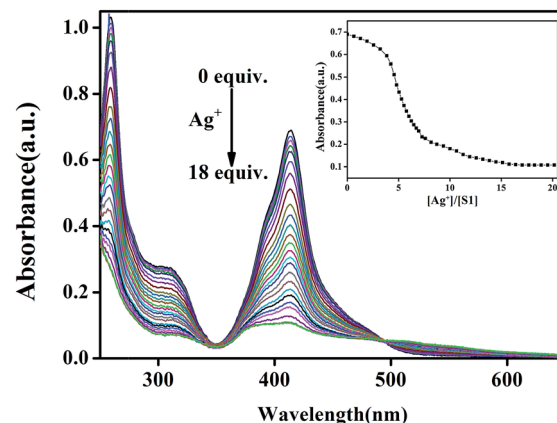


Fig. 4 Absorbance spectra of **S1** (20 μM) in the presence of different concentrations of Ag^+ (0–18 equiv.) in the DMSO/ H_2O (v:v = 1 : 1, buffered with HEPES pH = 7.2) solution. Inset: A plot of the variation in absorbance as a function of concentration of Ag^+ in the range from 0 to 18 equiv.

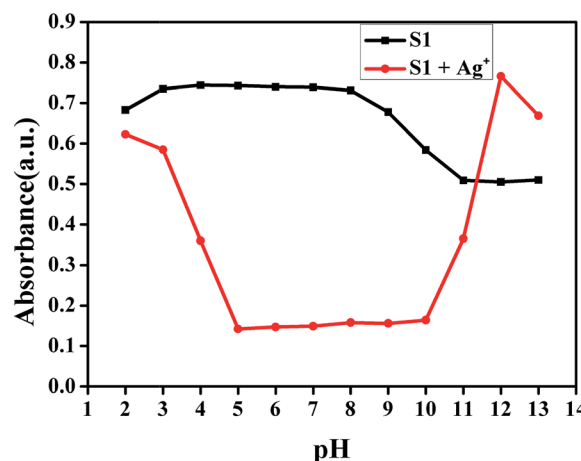


Fig. 5 pH response of **S1** and **S1**– Ag^+ in the DMSO/ H_2O (v:v = 1 : 1; buffered with HEPES) solution.

= 7.2) solution has been investigated. As shown in Fig. 5 and S7,† when the pH increased from 5 to 10, the absorbance and emission values remained almost unchanged. These results indicated that Ag^+ could be clearly detected *via* UV-vis and fluorescence spectral measurement using **S1** over a wide pH range.

The self-assembly and Ag^+ recognition mechanism of the **S1** sensor were investigated using concentration titration, ESI-MS, ^1H NMR titration, and DFT methods. To establish the supramolecular self-assembly mechanism, we performed the concentration titration in the absence and presence of different concentrations of **S1**. As shown in Fig. 6, the fluorescence intensity of the peak at 547 nm sharply increased with the addition of **S1**; this suggested that with the gradual increase in the concentration of **S1**, the self-assembled system was formed by π – π stacking and H-bonding in the aqueous solution. Moreover, the XRD spectra (Fig. S8†) support the



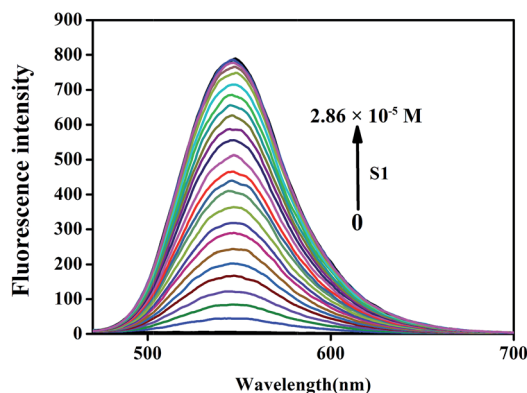


Fig. 6 Fluorescence spectra in the presence of different concentrations of **S1** (0 – 2.86×10^{-5} M) in the DMSO/H₂O (v:v = 1 : 1, buffered with HEPES pH = 7.2) solution.

abovementioned competitive mechanism. The free sensor **S1** exhibits a peak at $2\theta = 23.5^\circ$ corresponding to the d spacing of $3.82 \text{ \AA}$, which suggests that π – π stacking exists. In addition, compared to **S1**–Ag, this peak disappeared; this is attributed to the coordination of Ag^+ with the **S1** sensor. Based on all the abovementioned tests, we confirmed that **IM** self-assembly formed **S1** via π – π stacking and H-bonding in the aqueous solution.

The reversibility of the sensor function was tested by the addition of I^- to the **S1**–Ag complex. With the alternative addition of Ag^+ and I^- to the aqueous solution of the self-assembled **S1** sensor, the fluorescence emission of the tested solution exhibited alternating reviving and quenching processes (Fig. S9†). This off-on-off switching process was repeated several times. Therefore, the **S1** sensor can be considered a good ON–OFF fluorescence switch.

The 2 : 1 stoichiometry of the **S1** and Ag^+ complex was further confirmed by the results of the ESI-MS studies. The free probe shows a main peak at $m/z = 301.09$, corresponding to $[\text{IM} - \text{H}^+]$ (ESI-MS negative ion mode) (Fig. S14†), and the complex is confirmed by the appearance of a peak at $m/z = 731.12$, assignable to $[2\text{IM} + \text{Ag}^+ + \text{H}_2\text{O} + \text{H}^+]$ (ESI-MS positive ion mode) (Fig. S15†). The Job plot between **S1** and Ag^+ was implemented, demonstrating a 2 : 1 stoichiometry for **S1** and Ag^+ , as shown in Fig. S16.†

The FTIR spectra of the receptor **IM** and its complex with Ag^+ were obtained to examine the binding sites (Fig. S10†). In the FTIR spectrum of the free probe **IM**, the characteristic absorption band of **S1**– Ag^+ for N–H (3428 cm^{-1}) of the imidazole group was absent, and an obvious enhancement was observed for the C–N absorption bands from 1322 to 1421 cm^{-1} ; this confirmed the binding of **S1** with silver.

To gain more information about the complexation between **S1** and Ag^+ , we performed an ^1H NMR titration in the absence and presence of different concentrations of Ag^+ . As shown in Fig. S11,† the ^1H NMR titration of **S1** in DMSO- d_6 (10 mmol) with Ag^+ in D₂O (1 mol) established the stoichiometry and binding site towards the sensor response. Most interestingly, the N–H peak (H^1) of the imidazole group in **IM** disappeared after the

addition of 2.0 equiv. of Ag^+ to D₂O;¹⁵ this proved that **S1** was strongly bound with Ag^+ . Moreover, the signal of the hydrogen atoms in the phenazine (H^2 , H^4 , H^5 , H^7 , and H^8) and thiazole (H^3 , H^6 , and H^9) rings showed a significant downfield shift; this suggested that the intermolecular π – π stacking was destroyed after the addition of Ag^+ , and it further caused the deshielding for the hydrogen protons and resulted in downfield shifts.

To further investigate the binding behavior of **S1** with Ag^+ ions, we optimized the structure of **IM** and **IM**– Ag^+ complex by applying the B3LYP/6-31G(d, p) method in the gas phase using the Gaussian-09 computational code.¹⁶ The optimized structures of **IM** and **IM**– Ag^+ are shown in Fig. S12 and S13,† respectively. The short N34–H49...N23 distance of $2.041 \text{ \AA}$ indicates the presence of an intramolecular hydrogen bond in **IM**. Additionally, some of the relevant optimized geometric parameters are presented in Tables S2–S5.† The relevant occupied molecular orbital HOMO (-0.295 a.u.) of **IM** was localized mainly to the thiophene groups, whereas the unoccupied orbital LUMO (-0.227 a.u.) was localized over the phenazine and imidazole groups. The HOMO–LUMO energy gap (DE) for **IM** was found to be 0.068 a.u. (Fig. 7). On the other hand, upon interaction with Ag^+ , the HOMO (-0.291 a.u.) and LUMO (-0.279 a.u.) were almost localized on the entire molecule, where the electron density was polarized toward the whole molecule. The HOMO–LUMO energy (DE) for **IM** with Ag^+ was found to be 0.012 a.u. These results suggested that the fluorescence quenching was most probably attributed to the prevention of **IM** intramolecular charge transfer (ICT) by Ag^+ (Scheme 1).

To investigate the practical application of the chemosensor **S1**, test strips were prepared by immersing filter paper in a DMSO/water (v:v = 1 : 1, buffered with HEPES pH = 7.2) solution of **S1** and then drying them in air. The test strips containing **S1** were utilized to sense Ag^+ and other metal ions. As shown in Fig. 8, when Ag^+ and other ions were added to the test kits, an obvious color change was only observed for the Ag^+ solution under a 365 nm UV-lamp. The potential competitive

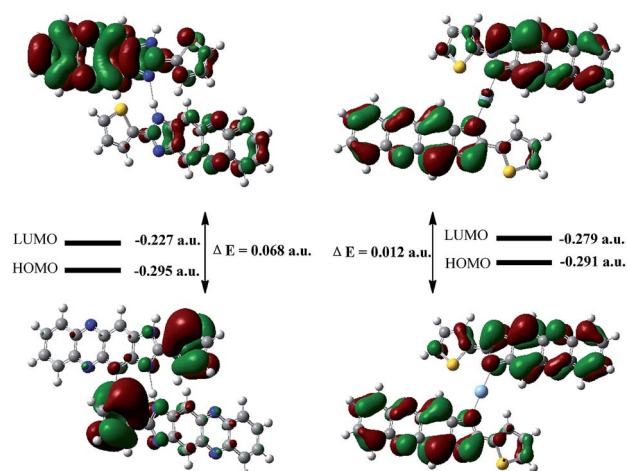


Fig. 7 DFT-optimized Frontier molecular orbitals and HOMO–LUMO energy gaps for **S1** (left) and **S1**– Ag^+ (right).



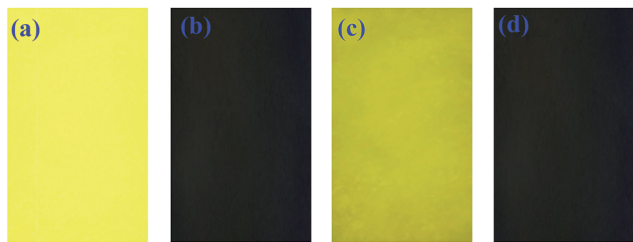


Fig. 8 Images of **S1** on test papers. (a) Only **S1**, (b) after immersion in DMSO/H₂O (v:v = 1 : 1, buffered with HEPES pH = 7.2) solution with Ag⁺, (c) after immersion in DMSO/H₂O (v:v = 1 : 1, buffered with HEPES pH = 7.2) solution with others cations, and (d) after immersion in DMSO/H₂O (v:v = 1 : 1, buffered with HEPES pH = 7.2) solution with Ag⁺ and other ions under irradiation at 365 nm.

ions did not influence the detection of Ag⁺ by the test strips. Therefore, the test strips could conveniently detect Ag⁺ in the solution.

Conclusions

In summary, we developed a novel self-assembled supramolecular sensor **S1** that could recognize Ag⁺ over other metal ions with significant UV-vis absorption and fluorescence quenching changes with highly selective and sensitive properties based on the competitive mechanism of supramolecular self-assembly. The obvious color changes and pronounced ON-OFF-type fluorescence signaling behavior can be seen by the naked eyes. The limit of detection for Ag⁺ was as low as 8.18×10^{-9} M in an aqueous solution, which was lower than that in some previously reported research. We believe that the sensor will be regarded as an environmentally friendly material that can sense Ag⁺ in an aqueous solution and make a great contribution to the development of silver ion sensors.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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