# EES Catalysis

## Accepted Manuscript

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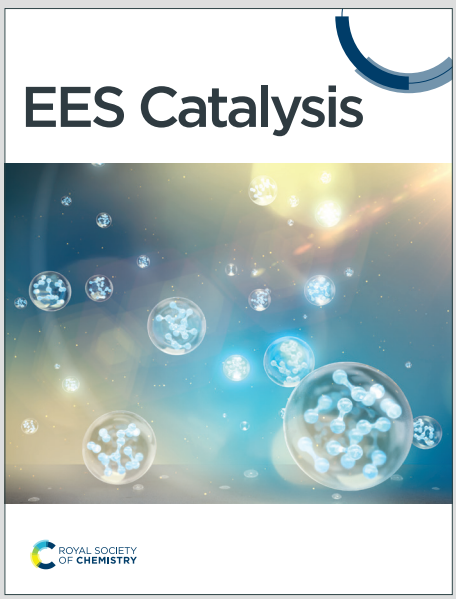




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Recent advancements in computational methodologies, particularly density functional theory (DFT) and materials informatics, have revolutionized catalyst design for oxygen reduction reactions (ORR). DFT, in conjunction with descriptor-based analysis and computational hydrogen electrode (CHE), not only aids in the rational design of catalysts but also offers insights into the influence of external factors such as solvent, field and pH effects on ORR activity. By enabling rational design strategies and accelerating the discovery of alternative materials, these approaches hold promise for overcoming limitations associated with traditional catalysts like platinum. Moreover, the integration of machine learning techniques and high-throughput DFT screening has facilitated the efficient generation and analysis of vast datasets, paving the way for the development of cost-effective and scalable ORR catalysts. This transformative landscape underscores the success of computational analysis in elucidating complex catalytic phenomena, offering unprecedented opportunities for the development of efficient and sustainable ORR catalysts, thus advancing the frontier of fuel cell technologies.



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# Revolutionizing ORR Catalyst Design through Computational Methodologies and Materials Informatics

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## Abstract

Computational approaches, such as density functional theory (DFT) in conjunction with descriptor-based analysis and computational hydrogen electrode, have enabled exploring the intricate interactions between catalyst surfaces and oxygen species allowing for the rational design of materials with optimized electronic structure and reactivity for oxygen reduction reaction (ORR). The identification of active sites and the tuning of catalyst compositions at the atomic scale have been facilitated by computational simulations, accelerating the discovery of promising ORR catalysts. In this contribution, the insights provided by the computational analysis to understand the fundamental reasons behind inherent ORR overpotentials in the experimental reported catalysts are discussed. Various strategies to overcome the limitations in ORR catalysis using computational design are discussed. Several alternative earth-abundant and cost-effective materials suggested by computational guidance to replace platinum-based catalysts are reviewed. The accuracy of DFT and the role of solvent and electrolyte pH are outlined based on the understanding provided by the computational insight. Finally, an overview of recent achievements in employing materials informatics to accelerate catalyst material discovery for ORR is provided. These computational advancements hold great promise for the development of efficient and cost-effective ORR catalysts, bringing us closer to realizing the full potential of fuel cells as efficient electrochemical energy conversion technologies.

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## 1. Introduction

There is a growing interest in creating renewable energy conversion technologies for automobile applications, such as fuel cells and metal-air batteries. Because of their remarkable prospects and ability to solve several complicated challenges that have long persisted, fuel cells stand out as the most promising future energy technology. In most fuel cells, the anode and cathode are separated by a membrane/electrolyte. At the cathode, oxygen is reduced to water, while a fuel, such as hydrogen, methanol, ethanol, or formic acid, is oxidized at the anode, releasing electrons and protons that pass through the external circuit and membrane to reach the cathode. Thus, fuel combines with oxygen without burning via a moderate electrochemical mechanism, with an optimal turnover frequency of $H_2$-$O_2$ of about 83% at 25°C.[1] In practice, however, low-temperature fuel cells do not achieve such efficiency, owing to the sluggish oxygen reduction process (ORR) at the cathode which results in a significant voltage loss. Platinum (Pt) has long been utilized as the most efficient ORR catalyst. However, there are several downsides to the extensive use of Pt-based catalysts in fuel cells. One of the most significant concerns is Pt scarcity and high cost, which causes scalability issues. Furthermore, even with the costly Pt catalyst, there is a significant overpotential associated with ORR (**Figure 1a**).[2,3] Furthermore, Pt-based catalysts have a low tolerance to methanol, which produces CO and blocks the active sites. As a result, the rational design of catalysts with even higher ORR activity than pure Pt is central to fuel-cell research. Various types of Pt-based catalysts have been largely examined over the past two decades. To boost the inherent activity of Pt-based catalysts, both alloying and morphological designs have been used. [4–6] The timeline in **Figure 1b** illustrates how Pt-based catalysts ORR performance has improved. As can be seen, the intrinsic activity of the catalysts has not changed, and there is still a 0.3 V overpotential that exists across these various Pt-based catalysts.

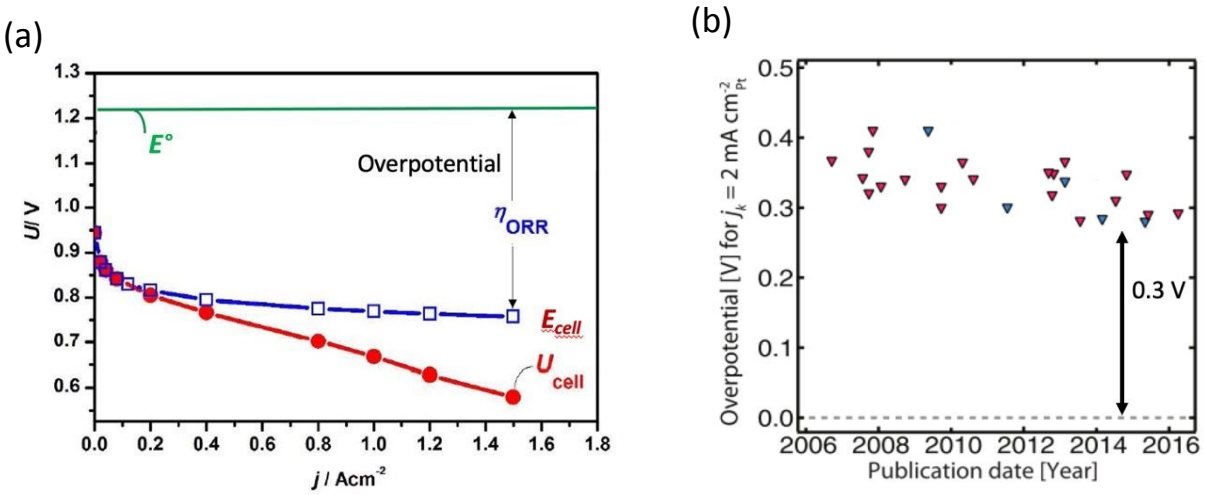


**Figure 1**. (a) Fuel cell electrochemical performance using Pt as catalyst. Copyright with permission from Ref. [7] (b) Timeline showing that the intrinsic activity of Pt-based catalysts has not been improved. Copyright with permission from Ref. [8].

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Computational approaches, such as density functional theory (DFT), in conjunction with descriptor-based analysis and computational hydrogen electrode (CHE) model suggested by Nørskov et *al*.[9], have played a pivotal role in elucidating the underlying mechanisms of ORR, understanding experimental results, and predicting novel catalyst materials with enhanced activity and stability. Herein, we first overview how descriptor-based analysis in conjunction with the CHE model can be leveraged to understand the fundamentals behind the ORR catalysis. Then, we show how this understanding can help to explain the large ORR overpotential for the studied classes of materials and guide the design of novel materials. Lastly, we show how these tools can be used to understand the ORR activity and stability of cost-effective and earth abundant materials such as transition metal nitrides and transition metal oxides. We should highlight that the examples mentioned in this contribution only focus on computational advances in ORR catalysis, and that they are only a sample of the vast array of literature on this subject. [10–12] Thus, this contribution is not intended to be exhaustive.

## 2. Fundamentals Behind ORR Catalysis

The past decade has witnessed significant progress in the computational understanding and design of catalysts for the ORR. Computational frameworks based on the surface science approach[9] and DFT calculations have played a pivotal role in elucidating the underlying mechanisms of ORR and predicting novel catalyst materials with enhanced activity and stability. These approaches have enabled direct theoretical investigations of reaction mechanisms on various model systems. Two-electron or four-electron oxygen reduction pathways could occur, depending on the catalyst properties. The two-electron pathway is known as partial reduction resulting in hydrogen peroxide (eq. 1) with OOH* as the only intermediate (eq. 2). The full reduction of oxygen allows the four-electron reduction, resulting in water (eq. 3). Herein, we only focus on the four-electron ORR which is of interest for emerging renewable energy technologies such as fuel cells, and rechargeable metal-air batteries. As mentioned above, currently, the sluggish oxygen reduction process hinders the efficiency of these technologies. Platinum (Pt) has long been utilized as the most efficient ORR catalyst, but it faces problems regarding scarcity and high-cost leading to scalability issues. Furthermore, even with the costly Pt catalyst, there is a significant overpotential associated with ORR (Figure 1a). To establish computational understanding for ORR catalysis, we need to simulate the reaction mechanism. Associative and dissociative mechanisms have been proposed for the four-electron ORR. The associative mechanism (eq. 4) involves three different intermediates, namely OOH*, O*, and OH* while dissociative one involves O*, and OH* (eq. 5).

Two-electron ORR:

$$O_2 + 2(H^+ + e^-) \rightarrow H_2O_2 \qquad E° = 0.70\ V \qquad (1)$$

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$$O_2 + (H^+ + e^-) \rightarrow OOH^* + (H^+ + e^-) \rightarrow H_2O_2 \quad (2)$$

Four-electron ORR:

$$O_2 + 4(H^+ + e^-) \rightarrow 2H_2O \qquad E° = 1.23\ V \quad (3)$$

Associative mechanism:

$$O_2 + (H^+ + e^-) \rightarrow OOH^* + (H^+ + e^-) \rightarrow O^* + (H^+ + e^-) \rightarrow OH^* + (H^+ + e^-) \rightarrow 2H_2O \quad (4)$$

Dissociative mechanism:

$$1/2O_2 + (H^+ + e^-) \rightarrow O^* + (H^+ + e^-) \rightarrow OH^* + (H^+ + e^-) \rightarrow H_2O \quad (5)$$

The associative mechanism is the commonly accepted mechanism for four-electron ORR. Taking this mechanism, the overall ORR catalytic activity is determined by the adsorption free energies of the reaction intermediates (OH*, O* and OOH*) to the catalyst surface. These adsorption free energies can be calculated using DFT calculations in conjunction with the CHE model, water stabilization[13], electric field effects,[14] and entropic corrections.[6] The CHE assumes the chemical potential of proton-electron pair is equal to the gas-phase $H_2$. The electrode potential is considered by shifting the electron energy by *–eU* where *e* and *U* are the elementary charge and the electrode potential, respectively.[9] These calculations can be used to build free energy diagrams along the reaction pathway (**Figure 2a-b**). **Figure 2a** displays the free energy diagrams for both the two- and the four-electron oxygen reduction reactions on ideal catalysts[15] at U= 0 V. This demonstrates a hypothetical case where the reaction is running by short-circuiting the cell and all the steps are strongly exothermic. The maximum potential allowed by thermodynamics is obtained by shifting the chemical potential of the electrons at the equilibrium potential of U= 0.70 V and U= 1.23 V for the two- and the four-electron ORR mechanisms, respectively.[9] In the ideal catalyst, all the steps in the free energy diagram are equivalent and take exactly as much as the equilibrium potential; hence they become thermoneutral at the equilibrium potential (green and magenta for four- and two-electron in **Figure 2a**). In practice, as in the case of the Pt surface, however, these steps are not equivalent to the equilibrium potential, and the potential at which the electrochemical reaction occurs is less than the thermodynamic limit (**Figure 2b**). This potential termed as thermodynamic limiting potential ($U_L$) can be extracted from the free energy diagram and is defined as the highest potential at which all the reaction steps are downhill in free energy. The overpotential is determined by the difference between the $U_L$ and the equilibrium potential.[15] The calculated $U_L$ for Pt (111) is 0.75 V, resulting in an overpotential of 0.48 V (**Figure 2b**).

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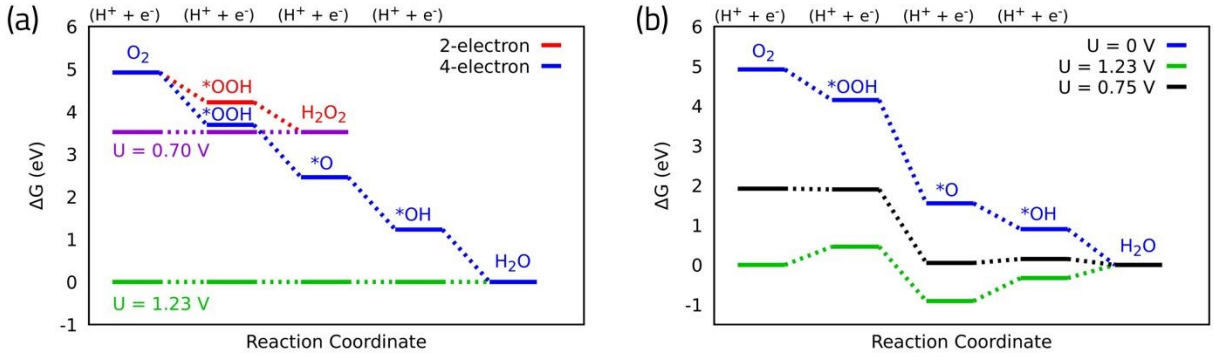


**Figure 2.** (a) Free energy diagram for oxygen reduction on ideal catalyst for two- and four-electron oxygen reduction. (b) Free energy diagram for oxygen reduction on Pt(111) associative mechanism. Data for Pt(111) are adapted from Ref. [15].

**Figure 3** displays the scaling relationship for the adsorption energies of the OH* and OOH* on a wide variety of catalysts structures. Based on this analysis, it has been demonstrated that the correlations between binding energies of different intermediates are observed on any catalyst surface (black dashed line in **Figure 3a**).[16–21] It has been shown that these correlations limit tuning the binding energy of one intermediate without affecting the binding energy of the other on any given catalyst surfaces. The green dashed line in **Figure 3a** displays the ideal correlation between OH* and OOH*. The star point displays the sweet spot for ORR catalyst, with four equal steps in the free energy diagram from $O_2$ to $H_2O$ (**Figure 3a**). The fact that the variety of known and examined materials are far from the ideal line causes a large negative impact on the activity by fixing the energy difference between OH* and OOH* around 3.2 ± 0.2 eV.[17–19] Because it takes two proton-coupled electron transfer steps to go from OOH* to OH* (reactions 2 and 3), a potential of at least 3.2 eV/2e = 1.6 V is needed to complete these reaction steps. However, the thermodynamic limit is 1.23 V, which means that even the best catalyst surfaces will have an overpotential of about 1.6 – 1.2 V = 0.4 V. [17–19] Taking Pt(111) surface for comparison, an ideal ORR catalyst should have an OH* binding energy approximately 0.4 eV weaker.[21]

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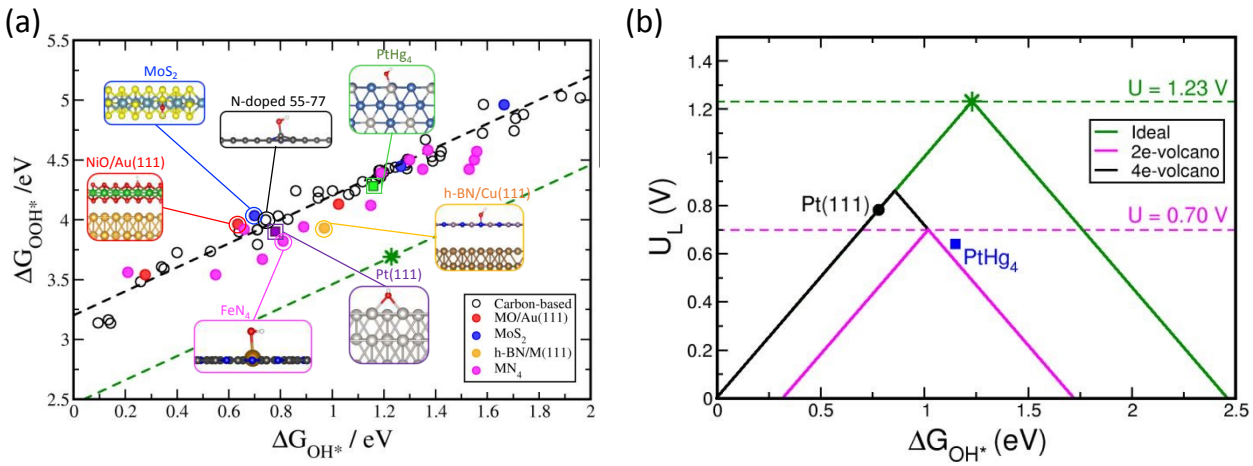


**Figure 3.** (a) Scaling relationship for the chemisorption energies of OH* and OOH* for various 2D materials. Pt(111) data shown for comparison in purple. The adsorption energies of OH and OOH are reported relative to liquid water and gas phase hydrogen using $H_2O(l) + * \rightarrow OH^* + 1/2H_2(g)$ and $2H_2O(l) + * \rightarrow OOH^* + 3/2H_2(g)$, respectively. Color code: C (gray), N (blue), O (red), H (white), B (pink), S (yellow), Mo (cyan), Cu (brown), Ni (green), Au (orange), and Pt (silver). (b) Ideal and conventional four-electron volcano plots in green and black solid lines, respectively. The magenta solid line displays the two-electron volcano plot. Green and magenta dashed lines display the equilibrium potentials for four- and two-electron processes, respectively.

Using the Sabatier principle, the activity of the catalyst surface can be related to the binding energies of the adsorbates as descriptors. This leads to a volcano shape plot, with the optimal catalyst marked a balance between not too weak, nor too strong binding of the adsorbates (**Figure 3b**).[19,22] Although any of the three adsorbates in the ORR can be selected as a descriptor of the activity since all of them scale with each other, typically, O* or OH* are used as descriptors. Taking OH* as an example, **Figure 3b** displays the ideal four-electron ORR activity volcano in green and the conventional four-electron volcano in black, which is different from the ideal volcano due to the scaling relations between OH* and OOH*. The negative impact of the scaling relation on the activity is clearly seen on the activity volcano, where the peak of the ideal four-electron volcano crosses the equilibrium potential as opposed to the conventional one. This indicates zero and ~0.4 V overpotentials for the ideal and conventional volcano plots, respectively. Therefore, the scaling relation limitation imposed by thermodynamics sets an upper limit to the maximum activity that can be obtained using the known classes of two-dimensional materials. On the plus side, the scaling relations have proved quite useful in explaining trends in oxygen reduction activity across various classes of catalyst materials.[9,22] It has also been successful in promoting the identification of new active catalysts for both two-[19,23] and four-electron oxygen[19] reduction reactions.

**Figure 3b** also displays the two-electron ORR activity volcano for reduction of oxygen to hydrogen peroxide in magenta. Since this reaction has only one intermediate (OOH*) the peak of the volcano crosses the equilibrium potential at 0.70 V. This in principle indicates that it is

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possible to find a catalyst with maximum activity if it binds that single intermediate with optimal strength, not too weak, nor too strong.

Platinum has long been used as the most efficient catalyst for ORR. Based on the above-mentioned thermodynamic analysis, it has been demonstrated that the thermodynamic sink for the oxygen reduction reaction on Pt(111) is the last step in the free energy diagram, the reduction of OH* to water (**Figure 2b**). In this thermodynamic picture, the ORR catalytic activity, relative to Pt(111), can be enhanced by weakening the binding energy of O* or OH*.[9,19] However, due to the scaling relationship between the binding energies of OH* and OOH*, beyond a certain weakening, it becomes thermodynamically uphill to activate $O_2$and form OOH*. It has also been demonstrated that decreasing the binding energy of O* or OH* in the order of 0.1-0.2 eV weaker than Pt enhances the activity.[9,19] This key understanding based on purely thermodynamic analysis paved the road for discovering new catalysts with enhanced ORR activity.[19] Using DFT calculations, a wide range of Pt alloys have been studied for identifying the best composition with 0.1-0.2 eV weaker binding energy than Pt.[19] Alloys of Pt and transition metals such as Co, Ni, Y and Sc were found to bind OH* weaker and showed enhanced ORR activity relative to pure Pt.[19] Both electronic and geometric effects play roles in tailoring the surface binding energy towards the favored regime. During past years, enormous efforts have been made to experimentally synthesize different Pt alloys and benchmark their activities.[6,12,21,24–29] Different design strategies have been investigated to control the structural composition and maximize the efficiency of the Pt alloy catalysts.[24,30,31] In line with the computational efforts, Xin et *al.*[30] proposed a model that relates the adsorption energies to the accessible physical properties of the metal element that forms alloy such as electronegativity, atomic radius, and spatial extent of valence orbitals.[30] Based on this model, the chemical environment of the Pt atom sites can be related to the local chemical reactivity. The accuracy of this model was further verified by DFT calculations.[30]

In addition to Pt alloy catalysts, a wide range of the other bimetallic alloys have also been studied to identify catalysts with enhanced ORR activity.[32,33] DFT calculations have also been used for mapping out the ORR activity of transition metal oxides,[34–36] carbon-based materials[35] and to less extent on transition metal nitrides[37] and sulfides.[38] We will delve into some of these findings in the forthcoming **Section 4**.

The DFT calculations have also been valuable in understanding the important role of the electrochemically most stable coverage of the ORR intermediates on the catalyst surface, as function of pH and potential.[39] This is particularly important for the catalyst surfaces with strong oxygen binding energy such as Ni where it has been shown that the activity drastically changes by including oxygen covered surface.[39] Using DFT calculations it has also been demonstrated that at low ORR potentials, the Pt surface is covered by the half-dissociated water layer and the O-O bond dissociation is clearly related to the local chemical environment.[13] Based on DFT and atomistic model for the charged solid–electrolyte interface[40], the barriers for proton transfer to the adsorbates on the Pt surface and barriers for proton transport within the water layer have

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been estimated to be negligible.[13] This, in turn, has enabled the modeling ORR through thermodynamic analysis alone, as a valid premise.

## 3. Guiding the Design of Next Generation ORR Catalysts

As mentioned previously, the activity of various examined catalyst surfaces is limited due to the scaling relation imposed by the OH* and OOH* correlations. Taking the associated ORR mechanism, the activity volcano is limited by two steps: OH* removal and OOH* formation on the surface. To deviate from the known scaling line and approach the ideal case, weaker OH* and stronger OOH* binding to the surface is desirable. Yet, achieving this goal has proven exceedingly challenging due to the scaling relation between OH* and OOH*, which renders it difficult to stabilize one oxygen intermediate while simultaneously destabilizing the other on a single catalyst surface.

In the past years, several strategies have been devised to circumvent the scaling relation and hence increase the ORR activity.[17,19,20,39–50] One strategy proposed by Siahrostami et *al.*,[37] is to divide the four-electron oxygen reduction reaction into two separate two-electron reduction reactions, i.e., 1) two-electron reduction of $O_2$ to $H_2O_2$ and 2) two-electron reduction of $H_2O_2$ to $H_2O$. They suggested that these reactions should be facilitated by separate catalyst materials: one highly active and selective for the partial reduction of $O_2$ to $H_2O_2$, and the other effective at converting $H_2O_2$ to $H_2O$ with both high activity and selectivity. Including a proton donor/acceptor site and the right combination of binding sites has been computationally analyzed and proven to improve the scaling relation. This idea was explored for several functional groups such as -COOH, -OH, and -$NH_2$ as proton-donor sites near proton-acceptor sites constructed of different transition metals coordinated with four nitrogen atoms (MN4). The combination of a nearby functional group such as -COOH and manganese metal coordinated with four nitrogen atoms (MnN4) was shown to make an ideal case for both ORR and OER.[17] Alternatively, it has been shown that molecular configurations like diporphyrin, which feature two metal sites, can effectively stabilize the OOH* intermediate in a dissociated state across both active sites, resembling an O* + OH* intermediate. Through systematic DFT calculations, the ORR activity of such molecular catalysts has been extensively examined, revealing a notable reliance on factors such as the intermetallic distance and the type of metals involved.[51]

Single atom catalysts (SACs) have also proven to be interesting for circumventing the ORR scaling relation due to their distinct properties surpassing those of traditional metal catalysts.[41] A comprehensive DFT investigation by considering more than 50 combinations of various metal single atoms embedded in various 2D substrates has unveiled distinctive electronic structure and geometric properties within SACs (**Figure 4a-b**).[41] These properties led to preferential stabilization of OOH* (inset in **Figure 4b**), consequently disrupting the scaling relationship between OOH* and OH* (**Figure 4a**), thereby enhancing ORR catalytic activity significantly via shifting the peak of activity volcano (**Figure 4b**).

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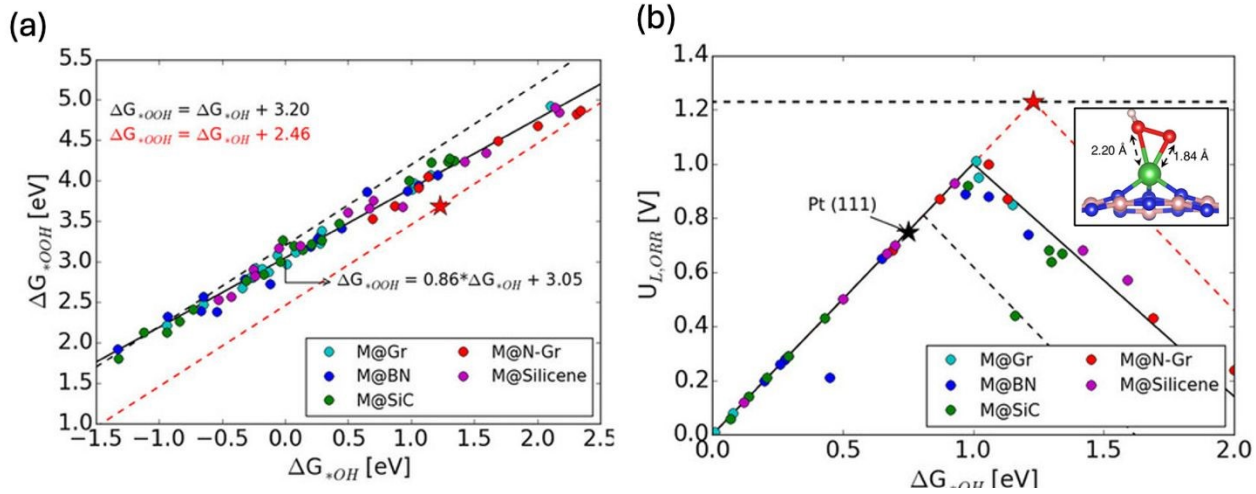


**Figure 4**: (a) Scaling relationship for the chemisorption energies of OH* and OOH* for various SACs. (b) Activity volcano plots for SACs (solid black), metal (111) (dashed black) and ideal catalyst (dashed red). Inset shows the bidentate adsorption stabilizes OOH* intermediate and results in overall improved activity. Copyright with permission from Ref. [41].

One of the interesting aspects of single-atom catalysts is their ability to interact with neighboring single-atom sites, which is crucial for fine-tuning their catalytic activity. It has been shown that adjacent single-atom catalysts, such as $FeN_3$ embedded in a graphene structure, exhibit communicative behavior upon the adsorption of small molecules like CO and $O_2$.[49] This behavior is attributed to long-range spin coupling. The O-O bond in $O_2$ is slightly more stretched when adsorbed on the second Fe site, and the adsorption energy changes by 0.18 eV. Interestingly, the Fe magnetic ordering shifts from ferromagnetic to antiferromagnetic in response to molecule adsorption.[49]

Dual-atom catalysts have also been recently investigated for the ORR. Xie, *et al*.[52] studied 144 different dual-atom configurations with M1-$N_4$ and M2-$N_4$ moieties embedded in graphene structures (where M1 and M2 are Mn, Fe, Co, Ni, Cu, Ru, Rh, Pd, Ag, Ir, Pt, or Au) using DFT calculations and machine learning techniques. They identified 13 dual-atom catalysts that lie at the peak of the ORR activity volcano, offering comparable performance to Pt but with the benefit of being composed of significantly cheaper and more abundant non-noble metals.

Another proposed strategy for overcoming the scaling relation in ORR is through confinement. Porous materials such as metal-organic frameworks (MOFs) have gained significant attention as potential ORR catalysts due to their unique properties in providing confined space.[51,53–56] MOFs are composed of metal ions or clusters coordinated to organic ligands, forming a three-dimensional porous structure. The combination of metal nodes and organic linkers results in a highly ordered and tunable framework (**Figure 5**). MOF structure provides numerous opportunities for confined spaces, potentially serving as an attractive platform for

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enhancing the ORR activity by altering the binding of OOH* and OH*. **Figure 5a and b** show an example of $Al_2(OH)_2$TCPP-Co which is composed of cobalt porphyrin catalytic metal center, aluminum hydroxide inorganic backbone and carboxylate phenyl benzene linker. Controlling the linker length allows adjusting the spacing between porphyrin centers. MOF structure, in particular $Al_2(OH)_2$TCPP-Co, has a lot of similarities with the structure of cytochrome C oxidase enzyme (CcO) (**Figure 5c**) which effectively oxidizes the catalytic cycle in respiration chain.[57] The enzyme design consists of separated Fe and Cu metals coordinated with N atoms in a porphyrin type structure.[57] This suggests that maximizing the efficiency of oxygen reduction reactions (ORR) could involve utilizing a metal-organic framework (MOF) structure featuring diverse metal catalytic centers within the cofacial porphyrin units. By employing a bimetallic model structure with an optimal distance between metal sites, illustrated schematically in **Figure 5d**, a strategy can be devised to target specific metal combinations that bind oxygenated species differently, thus maximizing ORR activity. In such a model, if one metal site exhibits oxophilic properties, it can readily bind oxygen species during the initial ORR cycle. These oxygen species, while acting as spectators, block the oxophilic site from participating in the ORR reaction but interact favorably with ORR intermediates forming on the active site (the other metal site). The most favorable outcome of this model structure is the interaction of spectators with the OOH* intermediate, stabilizing it via hydrogen bonding. Due to the size disparity between OOH* and OH* adsorbates, there is no hydrogen bonding stabilization from the interaction of spectators with OH* absorbed on the active site.

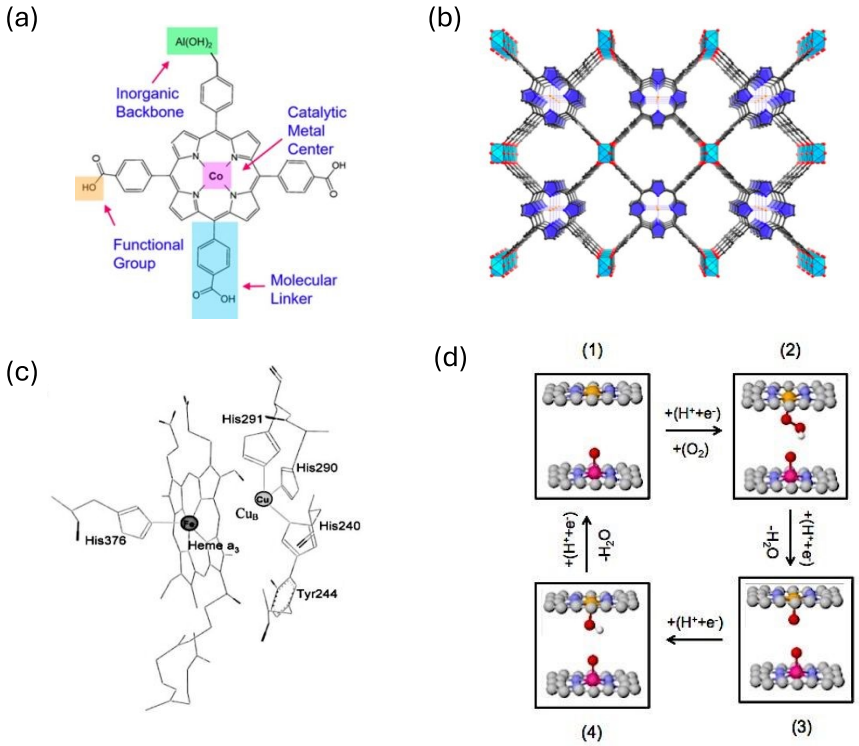

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**Figure 5.** (a) and (b) Coordination entity for constructing $Al_2(OH)_2TCPP$-Co MOF. Adapted from Ref. [54] with permission. (c) Cytochrome-C oxidase enzyme. Adapted from Ref. [57] with permission. (d) Illustration of the bio-mimic prototype catalyst capable of circumventing the scaling relations.

This hypothesis was examined in a recent computational study, affirming its effectiveness in controlling the scaling relations by varying the metal catalytic centers in $Al_2(OH)_2TCPP$-Co.[58] A bimetallic model structure of $Al_2(OH)_2TCPP$-Co-M, where M represents an oxophilic element, was scrutinized in this study as a platform to explore the effect of a third dimension in circumventing the scaling between OH* and OOH* and enhancing ORR activity. The bimetallic $Al_2(OH)_2TCPP$-Co-M, featuring metal centers approximately 7 Å apart within the cofacial porphyrin units (**Figure 6a**), was found to be optimal for the interaction of spectators with the OOH* intermediate. This study demonstrated that $Al_2(OH)_2TCPP$-Co-M, where M represents Fe, Cr, and Mn, effectively stabilizes the OOH* intermediate (Figure 6b), consequently altering the scaling relation (**Figure 6c**) and enhancing the ORR activity volcano relation (**Figure 6d**). Particularly notable is the combination of Fe-OH as a spectator and Cr as an active site, which provided the most effective confined local chemical environment for ORR, exhibiting the lowest calculated overpotential of 0.3 V (**Figure 6e-f**). [58]

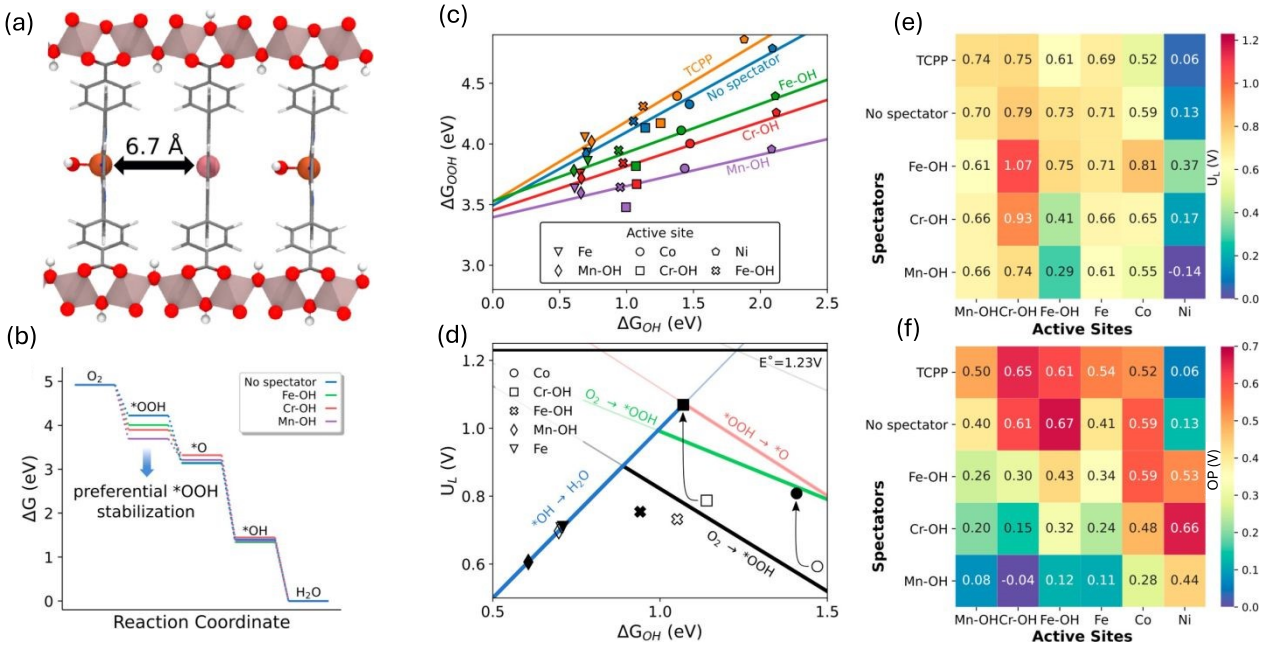


**Figure 6.** (a) Bimetallic model structure of $Al_2(OH)_2TCPP$-Co-M with an OH spectator siting on the oxophilic metal site. (b) Free energy diagram for ORR showing OOH* adsorbate free energy is most impacted by the spectator metal. (c) Scaling relations for the chemisorption energies of OOH* vs. OH* for various bimetallic MOFs. (d) ORR activity volcano plot showing a significant improvement using the confinement effect in bimetallic MOF (e) and (f) Calculated limiting potentials and overpotentials for different combination of the bimetallic MOFs. Adapted from Ref. [58] with permission.

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So far, we covered three strategies for circumventing the scaling relation between OOH* and OH*, including introducing oxophilic groups or proton-donor sites, using single atom catalysis, and fine-tuning the coordination environment through the third dimension. However, it is important to highlight additional strategies that have been suggested, which are well-covered in other review articles.[43,45] These include introducing p-block elements via heteroatom doping,[43,59] external field-assisted catalysis,[45] using ligand-modified catalysts,[43–45] strain effects induced by alloying including high entropy alloys,[59] spectator or co-adsorbed species,[44,54,58] geometry effects such as strain and surface curvature,[46,47] and dual-sites catalysts.[47,50,60] It is worth noting that geometry plays a major role in circumventing the scaling relationship among all the above strategies. One emerging approach is the concept of geometry-adaptive electrocatalysis, which holds significant potential in advancing catalytic performance. In a recent computational study by Cepitis et *al*.[44], the concept of geometry-adaptive electrocatalysis was examined to assess its ability to overcome the scaling relationship between OH* and OOH* Geometry-adaptive electrocatalysis represents a significant advancement in catalyst design. By allowing catalysts to dynamically adjust their geometry during reactions, this approach can potentially overcome limitations imposed by traditional scaling relations. This could lead to more efficient and effective ORR catalysts, where overpotential has been a persistent issue. [44] If experimentally validated, this concept could pave the way for developing next generation electrocatalysts with unprecedented performance, accelerating progress in energy conversion and storage technologies. The authors in demonstrated this concept using a model system of metal–nitrogen–carbon (M–N–C) catalysts, specifically the dual-atom site $2Co–N_4$ with variable curvature. Using DFT calculations, they demonstrated that altering the curvature of the catalyst can circumvent scaling through a dissociative mechanism, replacing the OOH* intermediate with O* on one site and OH* on another. This concept suggests the potential for discovering the ideal oxygen electrocatalyst.

Finally, we would like to emphasize that it is not merely about breaking scaling relationships; it is about breaking the right relationships in the right way to enhance the desired reactions without compromising catalytic efficiency.[61,62] This nuanced understanding underscores the need for targeted strategies that carefully consider the geometric and coordination environment factors influencing catalysis.

## 4. Earth-abundant ORR Catalysts

In addition to exploring novel chemistry and coordination environments to increase the intrinsic activity of the ORR catalyst, an interesting direction has emerged in the quest to identify catalysts that are both abundant on Earth and cost-effective. This endeavor aims to supersede Pt-based catalysts, paving the way for more scalable fuel cell technologies. Extensive research has been devoted to investigating various material classes.[19,63–66] Here, we focus on two notable categories: transition metal nitrides and transition metal oxides. However, the exploration does

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not end here; readers are encouraged to delve into additional review articles for a comprehensive understanding of alternative materials. [11,67–69]

**4.1. Transition metal Nitrides (TMNs) –** TMNs have lately received scholarly attention for their ability to selectively and efficiently catalyze the ORR.[63,70,71] These materials are sometimes referred to as "interstitial alloys", since the nitrogen atoms are incorporated into the interstitial sites of their parent metals.[72] An interesting characteristic of TMNs is their dependence on the ratio of the radii of the metal atom to that of nitrogen. Lattice expansion due to a radii mismatch can cause a metal *d*-band contraction and a higher density of states around the Fermi level, resulting in different catalytic characteristics than those observed in the parent metal. Furthermore, multiple stoichiometries can be produced from the same metal, allowing for an even broader range of catalytic characteristics and potential applications. [72] Herein we use DFT calculations in combination with descriptor-based analysis and CHE model to navigate through the vast chemical space of TMNs and identify promising ORR catalysts. We utilize the data published in Ref. [73] that employs high-throughput screening of 800 different TMNs from 22 metals listed in the Materials Project database (**Figure 7a-b**). This study shows that among the 800 TMNs, only 60 are thermodynamically stable with E above Hull ($E_{Hull}$) < 0.1 eV, indicating high stability with respect to decomposition at the same, fixed composition. Those 60 stable TMNs were considered for further investigation of $CO_2$ reduction reaction and an extensive database of 10,300 DFT calculations was reported. Among them, 6,430 data belong to unique OH* (3,220) and H* (3,210) adsorption energy calculations. We leverage this data to search for bifunctional catalysts for ORR and hydrogen oxidation reaction (HOR) for the cathode and anode of the fuel cell, respectively. **Figure 7c** illustrates the selection criteria employed to search for bifunctional ORR/HOR catalysts from the original dataset, taking Pt as the benchmark for both reactions. Using this screening scheme, 18 promising TMN candidates, composed of Fe, Co, Cr, Mn, Nb, and Ni, were obtained (**Figure 7d**). These 18 shortlisted candidates are earth-abundant and low-cost elements making them highly suitable catalysts to replace rare-earth and expensive Pt-based catalysts. Interestingly, Co, Cr, Fe, and Mn nitrides have already been reported as active catalysts for the ORR in fuel cells.[74,75] $Co_3N$ has been reported to show the best ORR performance, indicated by a mass activity of ~170 A $g^{-1}$ and an $E_{1/2}$ of 0.862 V vs. RHE, which is within 30 mV from the commercial benchmark Pt/C.[74] Computational modeling of CoN suggest a high activity with calculated limiting potential of 0.85 V.[70] $Nb_4N_5$, is a new promising ORR catalysts we identified using the theoretical data analysis in **Figure 7d**, which also finds applications as a superconductor material,[76] and a high-performance electrode material for supercapacitors[77] as well as for Li-S batteries.[78]

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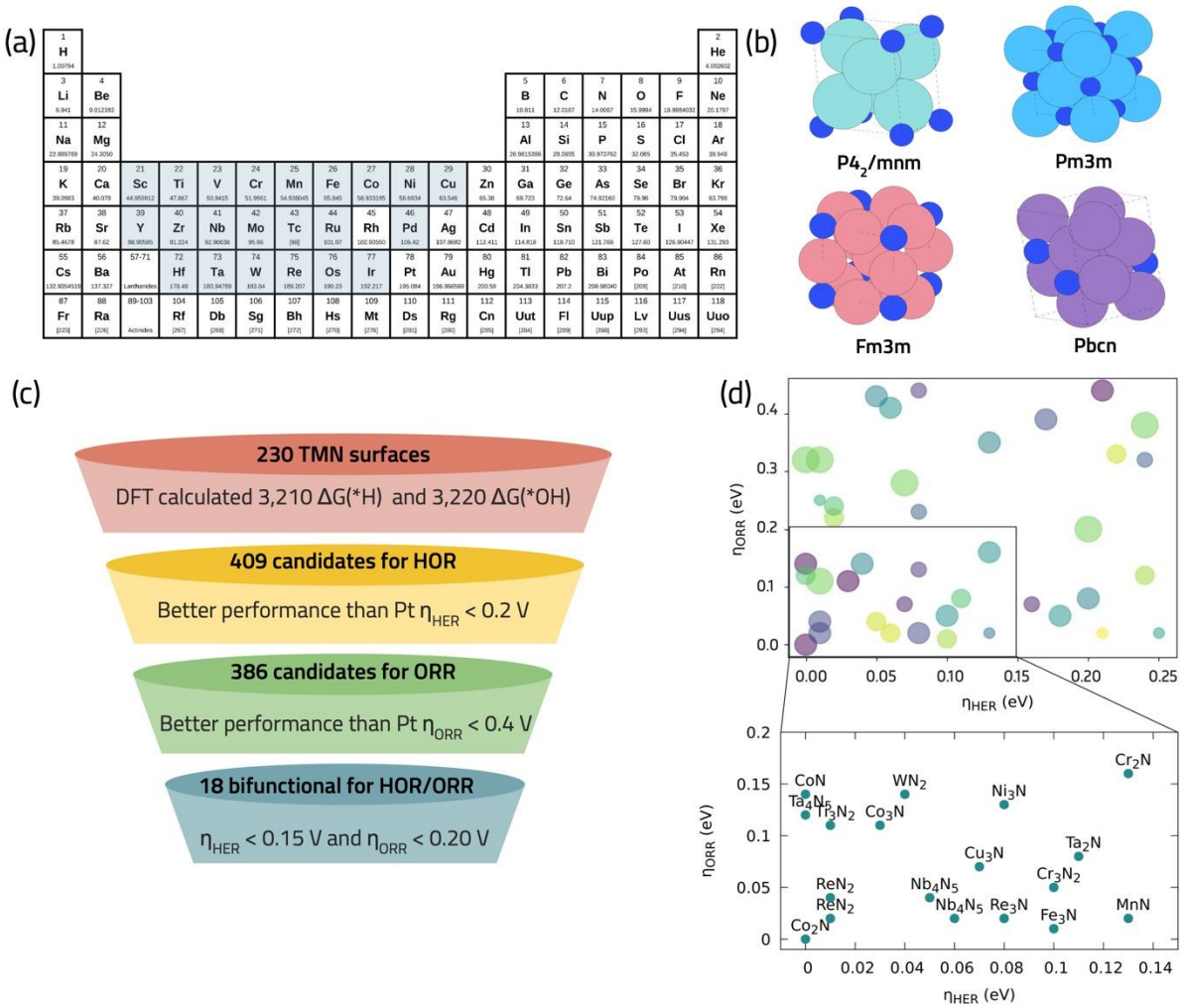


**Figure 7**. (a) Selected metals as base composition for TMNs from the Materials Project database. (b) examples of different bulk crystalline structures of TMNs studied in Ref. [73] with high thermodynamic stability. (c) Our selection criteria for identifying promising bifunctional TMNs for HOR/ORR. (d) Final promising list of bifunctional TMN candidates with low predicted overpotential for HOR/ORR.

It is interesting to note that, among this dataset, there are many promising candidates to promote ORR or HOR individually, as it is shown in a histogram distribution of the original data presented in **Figure 8a and b**. Namely, more than 600 TMN unique active sites can promote the HOR and more than 380 can promote the ORR taking *H and *OH as the HOR and ORR activity descriptors, respectively. **Figure 8c and d**, shows the free energy diagrams on several shortlisted bifunctional TMNs in **Figure 7d** that belong to first-row transition metals, and second/third row transition metals, respectively. We find that even though the obtained *OH adsorption energies for most surfaces agree with the data reported in, Ref. [73], the potential determining step does not come from the *OH formation in the examined catalysts. Instead, it arises from either the *O strong interaction with the surfaces, as in the case of CoN, MnN, and $Nb_4N_5$, with significantly negative Gibbs free energy values, or from the *OOH intermediate O-O bond scission, forming

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*O and *OH adsorbed in separate surface sites, as in the case of $Cr_3N_2$, $Fe_3N$, $Ni_3N$, $ReN_2$, and $Ta_4N_5$. These observations highlight the importance of considering the complete reaction mechanism to assertively determine the catalytic activity.

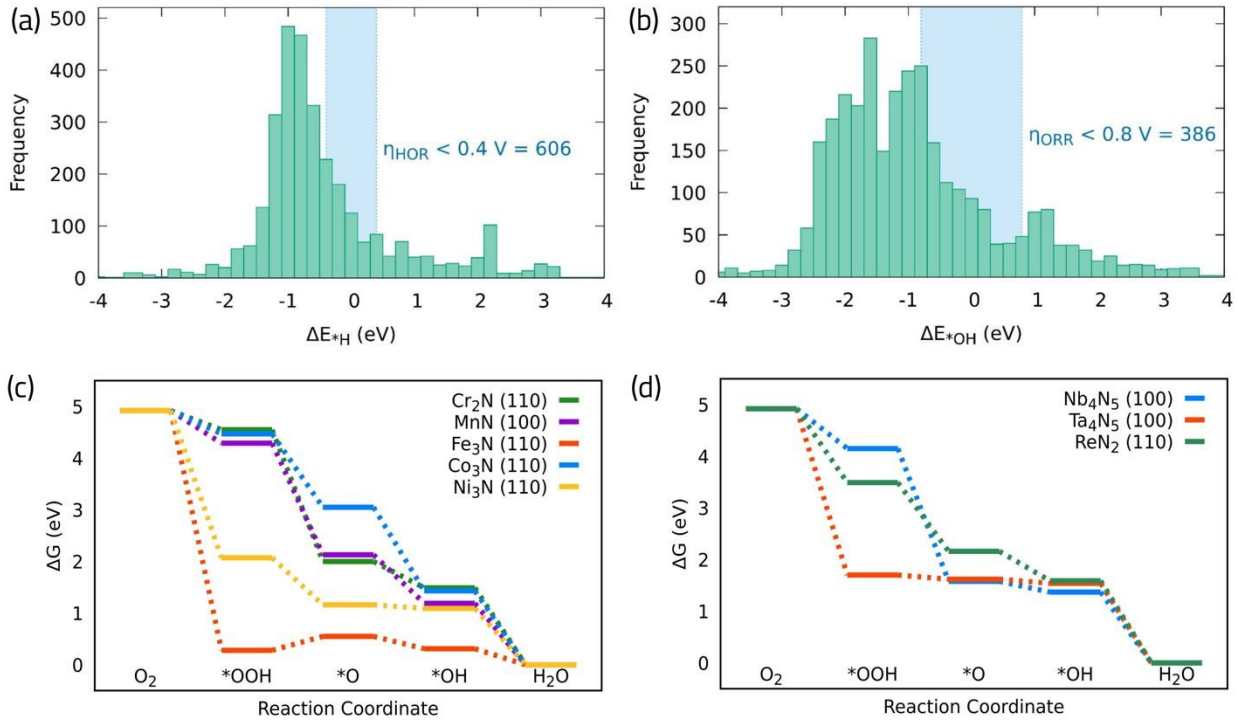


**Figure 8:** (a) and (b) Distribution of *H and*OH adsorption free energies across different TMNs studied in Ref. [73], respectively. (c) Calculated free energy diagram for ORR on different first row (d) and second/third row TMN surfaces. The corresponding data and structures can be found in the supplementary information.

The fact that oxygen binds strongly to the TMNs is in line with the previous DFT calculations [70] and experimental evidence[79,80], demonstrating the possibility of formation of oxide layer under ORR conditions. Although this can be taken as a negative characteristic, potentially indicating that TMNs are not stable under operating potentials for fuel cell applications,[81] Abroshan. H. *et al*.[70] demonstrated that the cobalt oxide layer formed on cobalt nitrides results in improving the ORR activity of TMN. Similarly, the formation of an oxide layer on nickel nitride[79] and molybdenum nitride[71] has been reported to improve not only its ORR activity, but also its stability.[71] A DFT calculated and precisely tailored Pourbaix diagram for $Ni_3N$, and $Ni_4N$ (as depicted in **Figure 9**), has revealed an insights into the expansion of oxide phases, such as NiOOH during the relevant pH and potentials of ORR compared to bulk Ni.[79] This analysis sheds further light on the intricate interplay between catalyst composition, electrochemical conditions, and oxide formation, crucial for understanding and optimizing catalyst performance in oxygen reduction reactions. This promoting effect of an oxide (or hydroxide) layer has also been

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experimentally observed in bimetallic nitrides, such as NiFeN and NiMoN,[82] vanadium nitrides,[83] and even for SACs/TMNs; where the hydroxyl coordination can improve the activity of the catalytic site by modulating the otherwise strong interaction with ORR intermediates. [84]

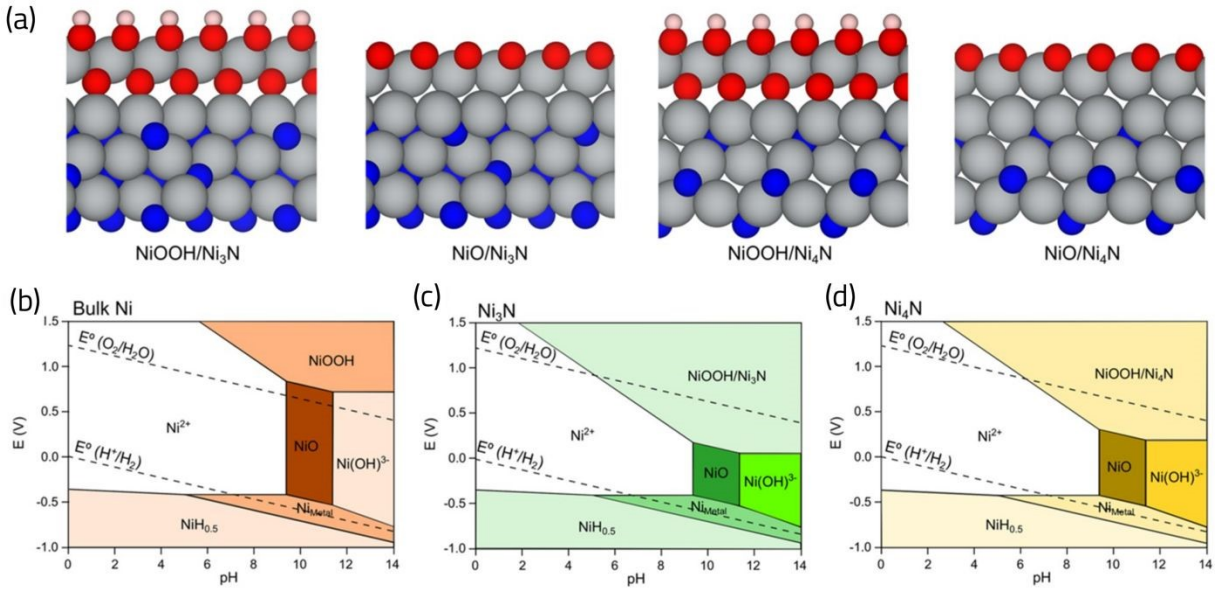


**Figure 9.** (**a)** Atomic structures showing formation of oxide (NiOOH and NiO) layers on $Ni_3N$ and $Ni_4N$ nitride surfaces. (b), (c) and (d) Constructed Pourbiax diagrams for Bulk Ni, $Ni_3N$ and $Ni_4N$, respectively. The extension of NiOOH phase is evident in (b) and (c) compared to bulk Ni, highlighting the propensity of nitride surfaces to develop oxide films under the relevant pH and potentials of ORR. Adapted from Ref. [79] with permission.

Similar results have been observed when transition metal sulfides are considered as ORR catalysts. Zhao, W. *et al*. [85] conducted comprehensive DFT calculations to investigate the oxygen reduction reaction (ORR) activity on $CoS_2$ catalyst.[85] Their findings suggest that sulfur is unlikely to serve as an active site for ORR. Additionally, they discovered that the undercoordinated Co metal site within $CoS_2$ demonstrates lower activity compared to the highly active undercoordinated Co metal sites within Co oxide films. This indicates that the ORR active sites are likely the oxide films formed atop $CoS_2$.

The sample studies mentioned above on ORR activity of transition metal nitrides and sulfides underscore the significance of accounting for the dynamic behaviors of readily available catalysts under ORR operating conditions and acknowledging the potential for oxide formation when constructing model structures. This relates to the importance of considering stability and the likelihood of catalyst surface restructuring during reaction conditions. To gain a deeper understanding of these effects, employing rigorous experimental techniques such as in-situ and operando measurements offers valuable means for exploring such phenomena.[67] Combined computational-experimental insights not only enhance the fundamental understanding of ORR

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mechanisms on earth-abundant catalysts but also hold promise for optimizing electronic structure of the catalysts to improve the efficiency of fuel cells and metal-air batteries.

**4.2. Transition metal oxides –** Transition metal oxides have garnered considerable interest recently owing to their abundance and low cost. These materials represent another class of compounds poised to revolutionize ORR catalysis, offering potential advantages in terms of both availability and durability compared to traditional Pt-based catalysts. One of the concerns, however, is their stability under ORR conditions in acidic media. A recent study by Wang *et a*l.[86] utilizes a large library of oxides reported in Materials Project database to screen the oxides that are stable under acidic pH values and relevant ORR potentials. This study suggests that the oxides of Sb, Ti, Sn, W, Mo, and Ge are acid stable presenting an opportunity to explore new classes of materials for ORR. Beyond stability, it is important to determine the catalytic activity of these materials to achieve a high ORR performance. [86]

A follow-up computational study,[87] examined antimonate (SbO)-based oxides functionalized with first-row transition metals ($MSb_2O_6$, M = Mn, Fe, Co, and Ni), considering different crystalline surfaces, to tune their activity and selectivity towards the four-electron ORR. The results indicated that catalytic activity was dependent on surface orientation, with the lowest theoretical overpotentials observed for (110) planes in most cases, and (100) for $MnSb_2O_6$. $MnSb_2O_6$ demonstrated the highest theoretical activity, which was corroborated by electrochemical experiments. These antimonates outperformed their pure oxide equivalents, $Mn_2O_3$(110), $Fe_2O_3$(001), $Co_3O_4$(100), and NiO(100), demonstrating the possibility for mixing and integrating different metals to modify catalytic activity. Experimental results on $MnSb_2O_6$ showed high selectivity towards the two-electron ORR can be achieved with further functionalization with Cr, Fe, and Ni. This study highlights the importance of combining different elements to achieve optimal performance on oxides.

Another particularly interesting oxide listed among the shortlisted candidates in Ref. [[86]], is oxides containing Sn. A common example is $SnO_2$ which is known to have a low ORR catalytic activity on its own.[88] Different strategies can be employed to improve its activity, such as different nanostructuring,[88] introducing oxygen vacancies[89] or heteroatom doping.[90] Herein, we investigate the effect of 18 different metal single atoms as potential dopants to improve the $SnO_2$(110) ORR activity, utilizing the DFT data reported in Ref. [[91]]. **Figure 10a** shows the stability analysis for these structures, calculated in terms of formation energy *vs.* dissolution potential. The formation energy indicates how favorable it is for a given element to occupy a vacant Sn site, taking the pristine oxide and the crystalline metal as references, with negative values (below 0 eV) indicating energetically favorable formation energy. Furthermore, the dissolution potential determines if the $SnO_2$-supported SACs will remain stable under ORR operating conditions.[91] All of the investigated SACs have greater dissolution potentials than 1.23 V, i.e., the standard redox potential for four-electron ORR, indicating their stability against dissolution. On the other hand,

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they all have negative formation energies when doped into $SnO_2$ substrate, meaning that they are thermodynamically stable single atoms. **Figure 10b** depicts the theoretical overpotential on both the Sn site and the doped site, offering a means to track the influence of each metal dopant on the ORR overpotential adjacent to the Sn site. **Figure 10c** shows the linear scaling[5] relationship between *OH and *OOH binding energies. This analysis reveals that most SACs have an improved ORR performance compared to pristine $SnO_2$, with the exceptions of Ca, Fe, and Ti. Sb, Nb, and Ta can significantly improve the ORR activity by lowering the overpotential even on the adjacent Sn site, which is highly desirable to reduce the required metal loading. The most promising SACs considering the doped metal as the active site are Ru and Pt, with low theoretical overpotentials around 0.35 and 0.45 V, respectively, placing them close to the top of the activity volcano plot (**Figure 10d**).

A similar approach has been reported by Mostaghimi et *al*. to tune the ORR catalytic activity of tantalum pentoxide ($Ta_2O_5$). [38] DFT calculations in conjunction with CHE and descriptor-based analysis was used to systematically investigated 22 transition metal atoms doped in $Ta_2O_5$(120). Their study revealed Pt, Rh, and Ir single atoms as the most promising catalytic active site, displaying improved ORR activity coupled with high stability. [38]

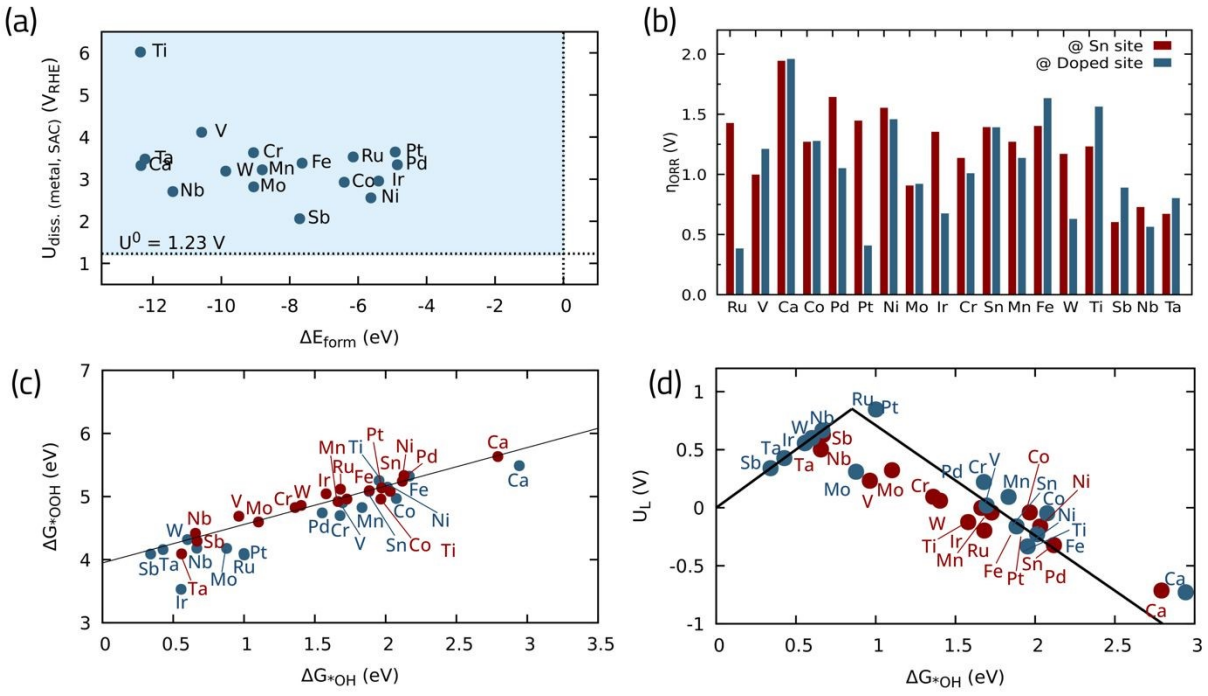


**Figure 10.** (a) Stability of single metal atoms in $SnO_2$ in terms of dissolution potential *vs*. formation energies, (b) theoretical overpotentials for the ORR with different doped-metals in $SnO_2$, (c) scaling relationship between ORR intermediates, and (d) activity volcano plot for the 4e-ORR. Data adapted from Ref. [91] with permission.

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**4.3. High Entropy Alloys (HEA) and High Entropy Oxides (HEO) –** High entropy alloys (HEAs) are metallic compositions that include at least five primary elements in about equal quantities. Unlike standard alloys, which often include one or two dominating elements and a small number of secondary elements, HEAs are designed to have numerous major elements with similar atomic sizes, resulting in a significant degree of disorder in the atomic structure, stabilizing the alloy as the entropic term dominates over the enthalpic one. Of note, not all HEAs are stabilized by entropy alone, and understanding their stability is an active research area. This unusual composition opens millions of new materials to be explored in the phase-space[92] with unique composition and electronic structure properties, and promising catalytic properties, making them increasingly interesting in ORR applications.[93] Löffler *et al*.[93] synthesized highly ordered $Pt_4FeCoCuNi$ nanoparticles with outstanding performance for both HER and ORR. The authors highlight the role of the structure ordering as a key parameter to tune the catalytic activity, as it was the case for another high entropy metallic catalyst, PtFeCoNiCuZn.[94] One of the other specific examples of high entropy alloys with high ORR activity is PdCuPtNiCo as reported experimentally by Chen *et al*.[95] Their synthesized nanoparticles had an average size of (10.4 ± 0.4) nm and were selective for the 4e-ORR, as indicated by the electron transfer of 3.95-3.97 observed in their experiments, and a half-wave potential (E1/2) of 0.83 V, very close to commercial Pt/C. Their stability tests showed a very small increase in E1/2 after 10,000 cycles.

Since the variety of binding sites of the different elements present in a HEA hinders computational modeling, it is crucial to have single-phase materials. Pittkowski *et al*.[96] investigated the PdCuPtNiCo using DFT calculations and demonstrated that the formation of HEA nanoparticles is governed by stochastic principles, meaning that their behavior and formation are influenced by random variables and probabilistic events rather than being entirely deterministic. They also showed that the inhibition of precursor mobility during the synthesis process favors the formation of a single phase in PdCuPtNiCo.[96]

The high entropy materials for ORR catalysis can go beyond metallic alloys and branch through other structures, such as perovskites, antiperovskites, oxides, and nitrides. Li, W. et *al*.[97] investigated lanthanum-based transition metal oxides ($LaTMO_3$) by substituting the B site with different transition metals (TM = Cr, Mn, Fe, Co, Ni) to form high-entropy oxides (HEOs) that tailor the *eg* occupancy by combining elements with different *d*-orbital electrons. $La(Cr_{0.2}Mn_{0.2}Fe_{0.2}Co_{0.2}Ni_{0.2})O_3$ was reported as the HEO with the best ORR performance, exhibiting an overpotential of 493 mV and a 1.7% decline in half-wave potential after 10,000 cycles. Despite these remarkable experimental results, the field still requires extensive exploration and development to utilize computational calculations as predictive models for such materials.

Currently, a lot of progress in the high entropy materials domain has been driven by experiments, and DFT calculations provide complementary analyses, often investigating limited combinations at a time. For instance, recent computational studies on the $LaMO_3$ structure

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employed DFT calculations to address the role of changing only the M site, tuning oxygen vacancies, or exploring different phases. [98–100] However, these efforts combined are still insufficient to predict the specific properties of an HEO such as the experimentally obtained $La(Cr_{0.2}Mn_{0.2}Fe_{0.2}Co_{0.2}Ni_{0.2})O_3$ [97] with acceptable accuracy. This concept was further elaborated in a recent work by Mints et *al*. [101] where a machine learning model was first trained using Gaussian Process (GP) regression using data from 350 synthesized nanoparticles via microwave solvothermal synthesis. After initial training and evaluation, the chemical space was reduced to IrOsPdPtRhRu. Each of these pure elements' oxides in the rutile structure and (110) surface was taken as references for incorporating the remaining atoms. At this step, ML techniques were again employed to reduce the number of DFT calculations required. The authors highlight the trade-offs from each part of this analysis and how computation and experiment complement each other. They show how information from experimental data is crucial but insufficient to determine whether the best performances arise from intrinsic activity or structural modification induced by the different elements in HEOs. ML models yield apparent correlations that might stem from data artifacts, necessitating careful analysis and filtering to ensure the maximum catalytic activity is not wrongly inflated. These models serve as necessary tools since DFT calculations cannot be performed for the entire sample space with all possible combinations of compositions and non-equivalent adsorption sites. Finally, DFT calculations do not address structural change effects but can predict intrinsic activity, and discrepancies among all three methods might occur.

In another example, Svane and Rossmeisl.[102] combined DFT calculations with other computational methods to investigate HEOs with rutile structures composed of Ru, Ti, Ir, Os, and Rh. Initially, 450 adsorption calculations were performed with DFT, and this dataset was used to train a linear fit model that considered different adsorption sites, summing up to 10,000 possibilities. This framework enabled the computational prediction of the optimal composition of an HEO to achieve the lowest overpotentials. Apart from HEA and HEO, antiperovskite nitrides such as ($InNCo_{2.7}Mn_{0.3}$) have been studied in combination with Pt nanoparticles as a bifunctional catalyst for OER and ORR.[103] This provide the opportunity to further explore high entropy TMNs where element modulation can be a very promising strategy to fine-tune different properties of these materials for ORR.[104]

In summary, modeling HEAs and HEOs with DFT poses a significant challenge, as obtaining any property in the HE materials requires numerous combinations of atomic distributions to accurately capture the random nature of these alloys. Trying to achieve that by brute force is unfeasible, as the number of simulations is too high, hence, methods to decrease the number of needed calculations, with some degree of approximation have been employed, such as virtual crystal approximation,[105] coherent potential approximation,[106] and alchemical potentials.[107] Machine learning algorithms, can be used to explore phase diagrams more efficiently, reducing the computational cost.[108] Pedersen, *et al*.[109] performed a Bayesian optimization of HEA compositions for ORR. Using DFT simulations, they estimated that approximately 50 experiments

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would be required to discover the optimal composition for quinary HEAs. Further details regarding this methodology will be provided in **Section 7**, where we delve into materials informatics for accelerating catalyst discovery for ORR.

## 5. Accuracy and Limitations of the DFT Analysis

**5.1. Choice of the exchange-correlation functional –** The accuracy of DFT calculations for predicting catalytic activities heavily depends on the choice of the exchange-correlation functional. The recent advancements in generalized gradient approximation (GGA)-based DFT functionals such as PBE, RPBE and BEEF-vdW allows accurate estimate of binding energy, with estimated errors ranging from 20 to 30 kJ/mol.[110] However, these functionals tend to overestimate the electronic and thermodynamic properties of bulk metal oxides. This overestimation primarily arises from a self-interaction error present in localized *d*- and *f*-electrons. To tackle this issue, the Hubbard *U* correction is commonly employed to address the strong onsite Coulomb interaction within oxides. However, determining suitable Hubbard *U* values is not straightforward and often requires benchmarking against experimental data. Cococcioni and Gironcoli[111,112] proposed a method to calculate Hubbard *U* values using a linear-response approach, providing a potential solution to this challenge. Other approaches can also be adopted, such as hybrid functionals[113] or Koopmans-compliant functionals.[114] Hybrid functionals combine local density approximation (LDA) and GGA functionals with part of the exchange from Hartree-Fock, to improve the description of exchange-correlation contribution and improve the accuracy. There are many types of hybrid functionals (e.g. Heyd-Scuseria-Ernzerhof (HSE) and Strongly Constrained and Appropriately Normed (SCAN)). Some have been shown to have a good performance for calculating binding energies. For instance, PBE0 and HSE06,[115] when compared to CCSD(T) benchmarks, show errors around 0.1 eV for calulating binding energies, where non-hybrid DFT methods such as PBE and RPBE have errors around 0.6 eV in some cases.[113] Koopmans-compliant functionals have also shown improvement in describing binding energies by eliminating single and many-particle self-interactions, known to adversely affect adsorption energies.[113,115] The main drawback of hybrid and Koopmans-compliant functionals is their computational cost, which are much higher than standard DFT or DFT+U approaches. Although, there are recent developments such as orbital-resolved DFT+U[116] could be an alternative to getting close to piecewise linearity without demanding much computational power.

Among the hybrid functionals, HSE functionals are popular because they can increase the accuracy of traditional DFT functionals such as PBE, especially for calculating band gaps in semiconductors. HSE incorporates a fraction of the exact Hartree-Fock exchange energy, which improves the description of electronic properties compared to standard DFT functionals. Apart from band gaps, HSE functionals are known to provide more accurate predictions of electronic structures and other properties of molecules, solids, and surfaces. For instance, Patel, A. M. *et al*.[92] reported that among 6 different functionals, the hybrid functional HSE06 with inclusion of

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dispersion corrections (D3)(BJ) provides the most accurate description of binding energies of ORR intermediates on a copper-modified covalent triazine framework. [113] In another report, double atom catalysts ($M_2N_6$) have been investigated for the ORR with HSE06 and PBE functionals.[117] The authors observed a similar trend with both, but employing HSE yielded a slightly better fit for the linear scaling relationships for reaction intermediates, and a shift in the spin magnetic moments of the adsorption site in respect to the peak of the volcano. Of note, HSE functionals are also computationally more expensive than standard DFT functionals, but slowly gaining more attention from the community, because they can yield more reliable results for systems such as metal oxides with strong magnetic properties, where the standard functionals fail to accurately describe the electronic structure.

The SCAN semi local density functional is another hybrid functional that was invented by Sun, J. *et al*.[118] It has been demonstrated to outperform GGA-PBE notably for a collection of 22 weak binding interaction energies, encompassing hydrogen bonds and van der Waals forces, with equilibrium values spanning from 0 to 20 kcal/mol. Similarly, it exhibits superior performance for lattice constants across a range of 46 hydrocarbons.[118] Additional enhancements to stability and performance have been achieved with rSCAN[119], which mitigates divergent behavior at low electron densities by regularizing SCAN's orbital indication function. Furthermore, the latest iteration, r$^2$SCAN, combines both approaches, incorporating previous adjustments to the isoorbital functions while reinstating certain original constraints. Kothakonda, M. *et al*.[120] showed that r$^2$SCAN provide more accurate enthalpies of formation closely aligning with experimental values, particularly when compared to PBE, for weakly bound solids. However, the authors acknowledge that PBE and PBEsol still offer greater accuracy for compounds containing transition metal elements in terms of formation enthalpy values. Nonetheless, challenges persist for hybrid functionals. Notably, it has been recently observed that r$^2$SCAN and r$^2$SCAN+U (with Hubbard *U* corrections) do not correctly predict the ground state electronic configurations of narrow band gap transition metal oxides.[121] Fundamental properties such as formation enthalpies, lattice parameters, band gaps and so on are still being tested to benchmark new hybrid functionals for different classes of materials. Of note, there is still considerable ground to cover in assessing whether these new functionals can offer improved descriptions of catalytic surfaces, along with their corresponding adsorption and reaction Gibbs free energy values, compared to those previously established in the literature.

**5.2. Accuracy of activity volcano plots –** Another crucial consideration when selecting functionals is to acknowledge that the accuracy of the volcano plot may be impacted by the choice of functional. Sargeant, et *al*. showed the importance of accounting for gas-phase deviations that depend on the exchange-correlation functional.[42] The authors note that the experimental value for the oxygen molecule is often taken as the reference to calculate the Gibbs free energy variations ($\Delta G_{exp}$ ($O_2$) = 4.92 eV), while the energy of this molecule calculated with a commonly

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used functional (GGA-PBE), would be, in fact, $\Delta G_{PBE}$ ($O_2$) = 4.46 eV. The difference between those values (4.46 − 4.92 = − 0.46 eV, in this case) is significant enough to shift the peaks of the volcano plots and yield different theoretical overpotential values. More specifically, the overpotentials will be overestimated, while the *OH adsorption energies will be underestimated. The mismatch was also observed for other functionals, such as RPBE, BEEF-vdW, and PW91.[42] This work shows the importance of semi-empirical corrections as a means of improving the accuracy of volcano plot predictions for ORR and OER catalysts. However, it is worth noting that when analyzing ORR trends, the potential concern about shifting the volcano plot may not be highly consequential. This is because, even if there's a systematic error affecting all the points, the trends in ORR activity can still be reliably captured.

## 6. Solvent, Field and pH effects

It is important to note that DFT calculations for calculating the adsorption energies of intermediates are performed in the gas phase. However, electrochemical reactions such as ORR deal with solvents and electrolytes. Thus, it is crucial to precisely consider the influence of water as well as the role of pH and cations in the electrolyte on the adsorption energies of reaction intermediates. Herein, we overview some of the computational efforts that have been conducted in the literature towards including such effects for ORR.

**6.1. Solvent effect –** Describing solvent effects to accurately represent solid-water interfaces using computational methods poses significant challenges for computational modeling due to the differences in bonding characteristics between water molecules as opposed to their interaction with the surface. Various computational approaches have been taken. Classical molecular dynamics, based on force fields, have been shown to effectively capture the statistical nature of liquid water but cannot address bond cleavage and formation.[122] While these simulations accurately reproduce the properties of water, they are unsuitable for investigating electrochemical processes, where new species form during reactions. Moreover, parameterizing force fields for each element and compound is a daunting task, especially when dealing with complex systems that include a catalyst surface (usually a metal, metal oxide, nitrides, or carbon-based solid), liquid water as the solvent, and various reaction intermediates (*O, *OH, *OOH). On the other hand, ab-initio molecular dynamics (AIMD) can describe both the trajectory of liquid water molecules and the formation of reaction intermediates.[122] However, AIMD is computationally demanding, making it impractical for significant surface models and high-throughput studies.[122]

Describing water-solid interfaces using explicit water molecules and DFT calculations is limited by the number of atoms in a simulation cell and periodic boundary conditions, which hinder accurate liquid water descriptions. Consequently, periodic models with very few explicit water molecules are used. For close-packed metals, metal oxides, and carbon-based materials, hexagonal monolayers or bilayers are typically studied. These models study the impact of water

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structure on the adsorption energies of ORR reaction intermediates on several close-packed metal surfaces, using hexagonal water bilayer structures.[13] These studies have revealed a sizable solvent stabilization (~ 0.3 eV) on the energetics of the ORR intermediates such as OH* and OOH*, but little to no effect on O*. Modeling the solvent effects on transition metal oxides and nitrides has been less extensively studied and reported using explicit water molecules. Thus, the influence of water molecules on ORR intermediates adsorbed on oxides remains unclear, primarily due to the challenge of obtaining accurate models of explicit water structures on various crystal structures and composition of oxides owing to the dynamic behavior of water molecules near the catalyst surfaces. Siahrostami, et *al*. [123], investigated the explicit water structure on three rutile oxide structures including $TiO_2$, $RuO_2$ and $IrO_2$. The goal was to describe the solvent effect on the oxygen evolution reaction (OER) intermediates. Their results show that the solvent has little impact on adsorption energies of the oxygen intermediates such as *O and *OH, except for the *OOH intermediate, where the solvent effect accounted for 0.4 eV of additional stabilization. Similarly, Gauthier, J. A. *et al*.[124] investigated the water structure on $IrO_2$ (110) surfaces by combining DFT calculations with global optimization (minima hopping) and explicitly modeling 1-3 ice-like water bilayers. The most stable water structures were composed of octagonal rings in this metallic oxide, which can be regarded as analogous to the hexagonal water structure formed on metallic surfaces, given the larger interatomic distances on $IrO_2$ (110) surfaces. When multiple water layers were stacked, pentagonal and heptagonal rings appeared between them, coordinated by hydrogen bonds. These two studies examining the impact of water structure on OER on the 110 surface of rutile oxides, bear significant resemblance to the ORR owing to analogous oxygen intermediates and can be used to estimate the solvent effect on other oxide surfaces.

For other classes of catalysts, such as transition metal nitrides, it is a common practice to use implicit solvent models to describe the solvent effect on the energetics of reaction intermediates as reported in the ORR study on CoN by Abroshan *et al*.[70] Implicit solvent models in which liquid water is modeled as a polarizable dielectric medium has been recently implemented in periodic DFT codes to describe solid-water interfaces.[122,125] These models are less computationally demanding and do not require exhaustive sampling of water molecules. However, as reported by Heenen *et al*.[126], the accuracy of the implicit solvent models in describing solvation effects at the solid-liquid interface remains unclear. This is because implicit solvent models usually do not account for the directional and steric interactions present with explicit solvent molecules. Thus, further research needs to be established to benchmark the implicit solvation models on the new classes of solid surfaces beyond metals.[126]

**6.2. Electrolyte and pH Effects -** Gaining computational insights into the pH effect on ORR activity and selectivity has been challenging, largely due to the disparity between the bulk property nature of pH and the surface properties addressed in DFT calculations, resulting in limited success. In a study by Kelly, S. *et al.*[127], pH was related to the field effect, with the argument that

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cations in an alkaline environment can induce a local field effect on adsorbed intermediates on the catalyst surface. For example, an alkaline environment (0.1 M KOH) has a high concentration of cations ($K^+$), in the electric double layer (EDL, region near the charged electrode surface where ions are present in solution). Hence, a local positive field is expected to be induced on the ORR intermediates adsorbed on the electrode surface in alkaline environment. They investigated the field effect on adsorbed oxygen species using sawtooth potential (**Figure 11a**) and showed that it is stronger for weak binding surfaces such as metallic gold. Then, they conducted a microkinetic modeling (**Figure 11b**) and concluded that for weak binding surfaces, kinetic barriers of the ORR depend on the electrolyte pH, with theoretically predicted overall barrier heights decreasing as pH values increase. [127]

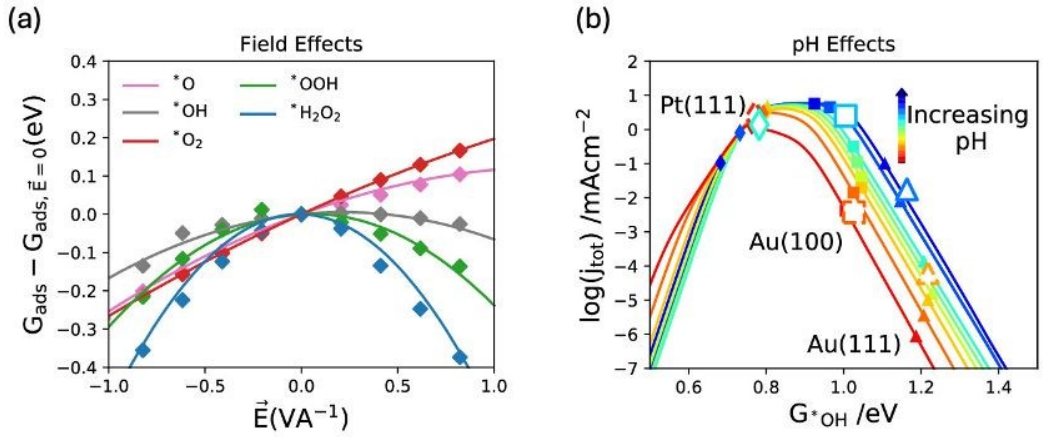


**Figure 11.** (a) Field effects on different oxygen reduction reaction intermediates and (b) activity volcano plot for the overall reaction at 0.9 V *vs* RHE under different fields. Adapted from Ref. [127] with permission.

The CHE model falls short in describing the field effect and consequently the pH effect, largely because the ORR intermediates binding energies are obtained from a zero-charge system and depend solely on the reversible hydrogen electrode (RHE) scale. The field effects, varies significantly with the pH, because of the equation below:

$$U_{SHE} = U_{RHE} + k_B T \ln(10) \times pH \qquad (6)$$

This relationship shows that for $U_{RHE}$ fixed at 1.00 V, $U_{SHE}$ varies within a range of 1.00 V to 0.17 V from pH = 0 to pH = 14, respectively, leading to different electric field strengths in the electric double layer.[127] The double layer contribution can be pivotal in an electrochemical process as it has recently been shown in the literature,[128] and it is not accounted for in standard DFT calculations. In addition to the electric double layer, variations in behavior under different pH levels and/or field strengths can also be affected by changes in the catalyst surface's potential of zero charge (PZC), resulting from its interaction with reaction intermediates.[129] These

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alterations can only be identified through explicit consideration of field effects in geometry optimization calculations of the adsorbed species. Furthermore, by uncoupling the contribution of the proton-electron pair, an evident dependency of the catalytic surface work function potential arises. This accounts for the observed magnitude difference in the additional stabilization obtained when field effects are considered between pure metal,[127] metallic oxides,[130] and carbon-based[66] surfaces, for example. A previous work from Duan and Henkelman[129] also treated the Au (100) surface as a case study using the double-reference method.[131] In this case, the solvent/metal interface is modeled by varying the number of electrons of the system. In order to compensate for the supercell charge difference, a uniform background counter charge is applied. Generally, by explicitly considering the field effects, the adsorption Gibbs free energy variation will fit a second order polynomial relationship, (**Figure 11a**), where $G_{ADS}$ is the adsorption Gibbs free energy at a given applied field ($E$), and the fitted coefficients α and μ are the intrinsic dipole moment and polarizability, respectively, as follows:

$$G_{ADS} = G_{ADS}^{PZC} + \mu E - \frac{\alpha}{2}E \qquad (7)$$

Taking pH and field effects into account has proven to be determinant to fill in the gaps between experimental observations and DFT calculations. For instance, only by modelling field effects it is possible to rationalize the difference in catalytic activity of low and high index transition metal oxides surfaces in acid and basic medium.[130] A recent study by Mao, X. *et al*.[132] incorporates the pH dependency and solvent contributions to provide a comprehensive screening through different Pd- and Pt-based alloys functionalized with a 1/3 monolayer of 30 different elements. The authors show that different alloys will have high catalytic performance under acidic (Ge/Pt, Hg/Pd, and Hg/Pt) and alkaline conditions (As/Pd, Cd/Pd, and Cu/Pd). Additionally, the generated database was employed to train a new theoretical predictive model that is mostly dependent on different surface parameters, with a cross-validation accuracy of 0.064 ± 0.007 eV on the test sets.

## 7. Materials Informatics for Expediting Catalyst Discovery for ORR

High throughput computational screening allows semi-automated simulations, where thousands of materials can be simulated and analyzed with little human intervention, allowing the creation of large databases.[133–135] With the large number of simulated materials, organized in databases, machine learning algorithms can be employed, for either classifications or prediction of materials properties.[73,136] This synergy has led to the emergence of a new area, materials informatics,[136–138] which is changing the paradigm on computational methodologies for materials science, with the catalysts design also being impacted by this new methodology. In addition, natural language processing techniques[139–141] are also being incorporated into the materials informatics frameworks, initially for data extraction from the literature by using regular expressions, then language models started to be employed for new applications of existing materials.[142]

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High-throughput DFT calculations have become a widespread practice to screen catalyst materials and generate datasets to train machine learning (ML) models. This approach has shown great promise for discovering highly active catalysts for ORR. Wang, Y. *et al.,*[64] conducted a high-throughput computational screening study on carbon-supported single metal atom catalysts. They carried out 180 simulations with various metals supported by different carbon structures and discovered that $\Delta G_{OH}$ serves as a reliable descriptor for ORR catalytic activity. They narrowed down the candidate catalysts to just five, which showed potential to outperform platinum. Sun, et *al*. [143] conducted a high-throughput screening for M-NC single-atom catalysts, encompassing various central metal atoms and environmental atoms. Their initial calculations produced 1,344 structures, of which they subjected 448 structures to DFT calculations using uniform distribution sampling. Utilizing this dataset, they trained an ML model to predict the overpotential, employing $\Delta G_{OH}$ as a descriptor for ORR catalytic activity. The authors employed 35 features independent of DFT results to train the ML model. They used feature importance methods to evaluate which factors contributed most to describing the target properties. For both the overpotential and the fourth step ΔG, the metal-related features accounted for approximately 88% of the contribution. These features included atomic radius, electronegativity, first ionization potential, electron affinity, the number of d electrons, the number of electrons in the outer shell, and the element's period in the periodic table. Other features related to atoms in layers 1 to 4 were also evaluated but contributed significantly less to the model. This led to the conclusion that if the central atom is not suitable for ORR, changing the surrounding atoms will have little effect on improving the catalytic process.

Of note, a model reliant on DFT calculations would not be suitable for predicting properties already obtained from DFT results. Their most effective model yielded an RMSE of 0.37 eV for $\Delta G_{OH}$ and 0.25 V for the overpotential, enabling them to infer these properties for the remaining 896 systems within the observed accuracy. To validate the accuracy of the ML-predicted results, the optimal candidates underwent further optimization using simulated DFT calculations to confirm their viability. To optimize their prediction, the authors trained another model, including a geometric feature (the sum of the bond lengths in the first layer) and the Bader charge. Performing all the DFT calculations to obtain these properties would negate the need for a model, as the targets would already be available. Therefore, the authors trained a model using the previous 35 features to predict these two new features, then inferred their ΔG and overpotentials. This model improved the RMSE by 0.02 eV.[143] They ended up reporting 30 highly active structures for ORR. Similarly, Chen, Y. *et al.*[144], applied ML algorithms trained with DFT results, for M-$N_4$-Gr/MXene heterojunction nanosheets. To accelerate their screening process, they applied regression algorithms to predict the overpotential, and their best model reached an RMSE of 0.10 eV. The authors also reached four suggestions of catalysts with low overpotentials.

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While the fusion of high-throughput DFT screening and ML algorithms presents significant opportunities for novel catalyst design, it is not devoid of limitations. High-throughput DFT calculations demand significant computational resources, which are not accessible to all researchers. On the other hand, the ML methods do not offer the same precision as DFT, and they are still not able to achieve the chemical accuracy of 1 kcal/mol. A major concern when applying ML techniques to aid catalyst discovery is that these often result in "black box" models, making it impossible to understand the underlying chemical and physical properties that contribute to good catalytic activity purely from ML selection. To address this challenge, explainable AI[145] has emerged as an alternative to combine the low cost and extensive data screening capability of ML models with the incorporation of theoretical and physical principles, particularly in material discovery. Coupling both high-throughput and machine learning (ML) in active learning process has proven to be a key. [146] Active learning algorithms can be implemented in ORR catalyst design to enhance the efficiency and effectiveness of discovering new catalysts. This approach involves iteratively training machine learning models on an initial dataset of known ORR catalysts, using strategies like uncertainty sampling to select the most informative samples from a pool of unlabeled data. These selected samples are then experimentally tested, and the results are added to the training set to refine the model. This cycle continues until the model achieves a satisfactory performance or the available resources are exhausted. Active learning helps prioritize experiments that provide the most value in terms of information gain, thereby speeding up the discovery process. The implementation of active learning for ORR seems to be in its early phase, with limited literature currently available on the topic. Zhang X. *et al*.[147], studied platinum-based alloys, and identified five promising candidates for ORR with low overpotentials after just three iterations of the active learning process. Of note, none of the features used for the ML model involved DFT simulation. Another active learning study, by Omidvar N. *et al*.[145], investigated Pt monolayer core-shell catalysts. A recent study by Omidvar N. et *al*. [145] employs active learning to investigate Pt monolayer core-shell catalysts using theory-infused neural network (TinNet) algorithms.[148] The authors explored approximately 17,000 candidates generated from roughly 1,500 thermodynamically stable bulk structures sourced from databases. They performed 8 iterations and successfully identified previously known ORR catalysts with high activity, thereby validating their methodology. A theory module based on the Newns-Anderson model Hamiltonians, which models chemisorption processes, was incorporated into the neural networks to provide model interpretability. By factoring in material costs within their selection process, they steered away from precious metals and identified 17 candidate materials that offer both cost-effectiveness and high ORR performance. These examples indicate the great potential of active learning methods in expediting catalyst discovery for ORR. However, the main challenge is to develop machine learning models that can provide reasonable performance with few calculations. These models do not need to be very precise since they are primarily used to suggest promising candidates, with precise results obtained through

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subsequent calculations. Nevertheless, the models must reliably estimate whether a catalyst will perform well or poorly. Additionally, there is a trade-off between exploring and exploiting the phase space. Focusing solely on simulation suggestions for the best-performing catalysts predicted by the model might lead to missing optimal regions in the phase space where better catalysts could be discovered. Therefore, when performing calculations for new suggestions, it is crucial to maintain a balance between evaluating both good and poor-performing catalysts to ensure comprehensive exploration of the phase space.

Current methodologies clearly result in a significant volume of data being generated in literature. Some of this data is organized into databases, however, most findings are dispersed across published papers, in PDF format, with XML format emerging more recently. To address this challenge, natural language processing (NLP) methods hold significant promise. They can leverage the wealth of existing data, derived from both experimental and theoretical studies. Surprisingly, NLP methodologies seem underutilized in the literature for the ORR studies, with most works focusing on $CO_2$RR. In the study by Suvarna, M. *et al*.[149] data collecting has been performed for ORR, OER, HER, and $CO_2$RR. They reported the most used metals precursors, carrier materials, and solvents, with Fe being the most investigated metal for ORR. For example, ZIF-8 frameworks were commonly used as precursors in preparing ORR catalysts. Hence, NLP methods are a powerful tool for obtaining data from literature, although its main limitation is a method to extract data from published literature. Accessing papers collectively is challenging, as not all journals permit such access or may impose additional fees, on top of the standard access fees paid by research institutions. Furthermore, the usual brute-force techniques are poor and demanding to obtain context-based information, hence, the large language models (LLMs) are playing a transformative role, as they are particularly good at capturing the context with long term relations. Another limitation is the cost to perform API queries with the best performing LLMs, although running LLMs locally is becoming feasible, as the performance of open weight models is improving fast.

Currently, there are only a few works that broadly apply NLP to include ORR results. Here, we discuss two insightful examples from the literature that contribute to advancements in ORR catalyst design. Qin, *et al*.[150] employed NLP for insights into the development of electrocatalysts. They gathered 604,933 abstracts from the literature and trained a named entity recognition (NER) model. The authors identified electrocatalysts and their performance, extracting 22,000 records of catalysts from their corpus. Although the data obtained was limited for ORR, they found that platinum remains the preferred material for ORR, despite many alternative materials being studied. These results differ from our findings, where we show a trend toward metal nitrides for ORR, as discussed in the next paragraph. Another NLP study by Muthukkumaran et *al*.[151] focused on discovering perovskite-based electrocatalysts. The authors retrained a SciBERT model with 1.74 million abstracts, using it to generate new perovskite compositions with synthetic embedding. To evaluate the proposed compositions, they applied various cosine

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similarity measures, including the similarity to "electrocatalyst" and comparisons between word embeddings obtained from the language model and those generated by the algorithms. The authors concluded that nickel-based perovskites are promising candidates for catalysts.

To illustrate the potential of the NLP techniques for ORR, we mined 34,515 abstracts from the Web of Science query "oxygen reduction reaction", encompassing all relevant studies on this topic (**Figure 12**). We further filtered the results, by keeping the abstracts that did not have expressions related to other catalytic processes, leading to the final number of literatures down to 20,237. This allowed us to narrow down to papers that are more focused on the ORR. Research on ORR began in the early 1990s, experiencing a growing trend since then. However, it was not until 2006 that there were more than 100 works, and the milestone of 1,000 papers was surpassed in 2014. By 2022, the number of works had reached 2,000. With the collected and filtered abstracts, we employed the large language model Llama 3 70b,[152] to extract which materials were used as catalysts in each work. The analysis was performed since 2006, as the previous years did not have enough papers to provide good statistics. **Figure 12a** plots all the materials mentioned in the abstracts as catalysts that had at least 0.6% of the total mentions for at least four years. Platinum prevails in publications as a state-of-the-art catalyst material. However, its mentions have been steadily decreasing over time, even when accounting for its potential inclusion as a catalyst benchmark, which could increase its occurrences. Examining the lower portion of **Figure 12a** reveals the trends of other materials. For instance, there is a declining trend in the share of works utilizing Pd, Au, Ag, Pt-Co, Cu, Pt-Ni, Ru, and others. For Fe, Co, and particularly Fe-N, the share of works is on the rise. To highlight emerging trends in materials, we have plotted those with increasing trends over the last four years and at least 0.08% of mentions in 2023, as depicted in **Figure 12b**. Many materials containing nitrogen are experiencing growth, notably Fe-N, as mentioned earlier, along with Co-N, Fe-NC, M-N (metal-nitrogen), g-$C_3N_4$ (graphitic carbon nitride), Mn-N, and Cu-N. Transition metals like Co and Fe are also experiencing an upward trajectory, albeit not as significantly compared to the initial years of analysis, given their already similar shares. $Fe_3C$ and Mn are also witnessing increases, albeit with lower total shares. There is a diversification in the materials under study, with platinum losing shares while others gain prominence. Thus, it is evident that NLP techniques can offer valuable insights from the analysis of entire literature sets, even when focusing solely on abstracts, which are readily accessible without paywalls. A more comprehensive study could delve into this dataset, offering a detailed examination of rising and declining trends in research for each material, given that many materials were excluded due to the applied filters. Furthermore, properties such as overpotentials, limiting potentials, or Faradaic efficiencies could be extracted, associated with materials, facilitated by recent advancements in LLMs, representing a promising avenue for future research. Access to full texts via mass paper download tools from publishers could significantly expand the scope of NLP work, although such tools are not readily and affordably accessible.

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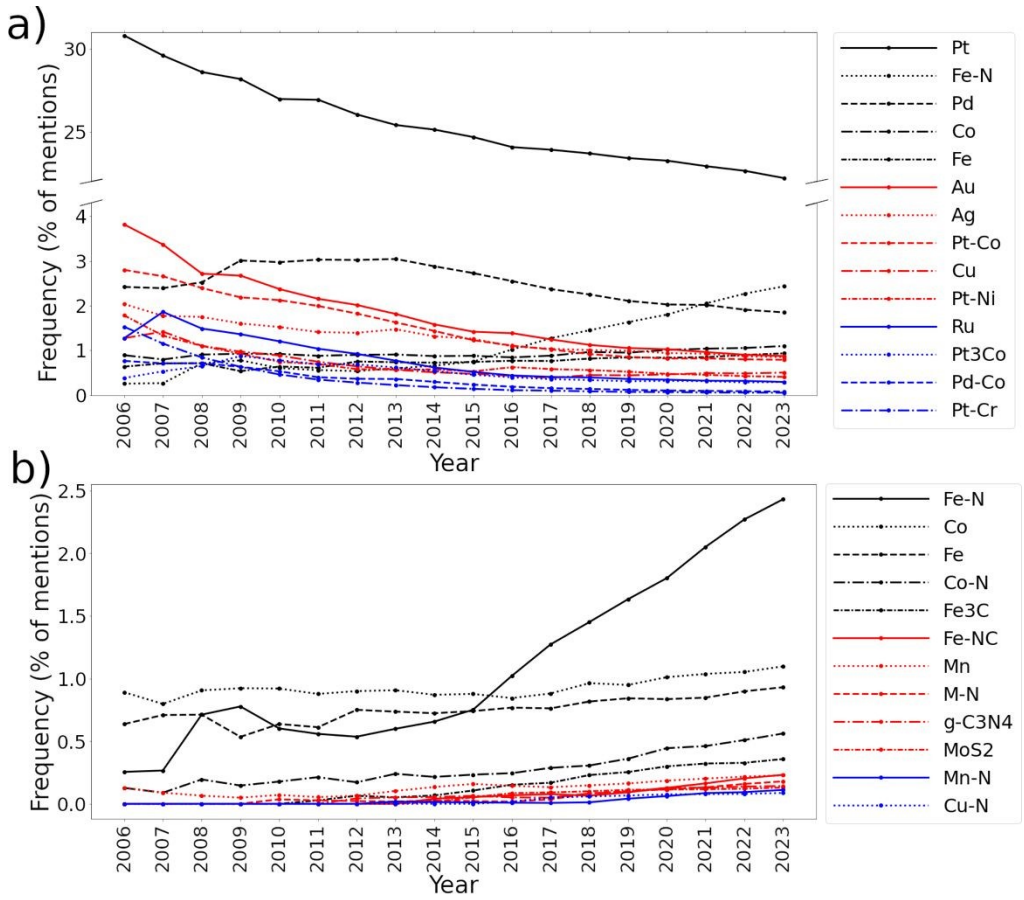


**Figure 12.** (a) Number of publications per year for all reported materials for ORR. (b) Number of publications per year for rising trends over the last four years.

## 8. Summary and Perspectives

Computational methodologies, including DFT in conjunction with descriptor-based analysis and CHE, have revolutionized the study of catalyst surfaces and their interactions with oxygen species. This approach enabled rational design of materials with optimized electronic structure and reactivity for ORR. By identifying active sites and fine-tuning catalyst compositions at the atomic level, computational simulations have expedited the discovery of promising ORR catalysts. The insights gained from computational analyses have shed light on the fundamental causes of ORR overpotentials observed in experimental catalysts. Various strategies have been explored to overcome limitations in ORR catalysis through computational design. Alternative earth-abundant and cost-effective materials such as transition metal nitrides and oxides have been recommended by computational guidance, providing opportunities for replacing costly platinum-based catalysts and solving the scalability issue of fuel cell technology. The accuracy of DFT calculations and the influence of solvent and electrolyte pH remain ongoing challenges to tackle using computational methodologies.

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In addition to unraveling atomistic-level information that contributes to understanding the origin of catalytic performance, computational methods also serve as predictive tools to guide experimental synthesis. Numerous examples in the literature demonstrate the effectiveness of computational screening in identifying novel ORR catalysts. These include the discovery of various Pt-alloys,[4,6,7,19] oxides[67] such as antimonates ($MSb_2O_6$, M = Mn, Fe, Co, and Ni),[87] and high-entropy alloys like $Pt_4FeCoCuNi$.[93] Recently Ir alloys ($Ir_3M$, where M represents 3d, 4d, and 5d transition metals) predicted as active ORR catalysts using DFT calculations, which were subsequently synthesized and verified through experimental testing.[60]

Recent advancements in employing materials informatics, including machine learning approaches like NLP, provide opportunities to accelerate catalyst material discovery for ORR. Furthermore, the application of high-throughput DFT screening techniques has enhanced this process. These methods enable the efficient analysis of vast datasets, facilitating the identification of trends and patterns that inform the design of novel ORR catalyst materials. Computational advancements in machine learning models, coupled with high-throughput DFT screening, hold immense promise for developing efficient and cost-effective ORR catalysts. Ultimately, these efforts aim to advance fuel cells as highly efficient electrochemical energy conversion technologies, bringing us closer to realizing their full potential.

As more data is generated with high-throughput screenings, the importance of databases to store these calculations cannot be overstated. These databases ensure that the results of calculations are accessible across different fields, particularly in experimental research and development. This is crucial not only to ensure the reproducibility of DFT calculations and to reduce unnecessary computational time and resources being employed for redundant analysis but, most importantly, to accelerate catalyst design. A recent milestone in this regard is the creation of Catalysis-Hub[135], a comprehensive database of reaction energies obtained from electronic structure calculations, with a user-friendly web interface, to which more than 130,000 calculations have been uploaded since its announcement.

**Author contributions:**

S.S. conceived the idea and structure of the review paper, wrote the first draft, and supervised L.E.B.L. in conducting the screening study on ORR activity of transition metal nitrides. J.M.A. performed the natural language processing. All authors collectively contributed content to enrich the paper and participated in revising and editing.

**Conflict of Interest:**

There are no conflicts to declare.

**Data Availability**

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The data supporting this article have been included as part of the Supplementary Information. This file contains all original data for the DFT calculations presented in the main manuscript **Figure 8**.

**Acknowledgments:**

S.S. and L.E.B.L. gratefully acknowledge the support from Natural Sciences and Engineering Research Council (NSERC) of Canada (Discovery Grant No. RGPIN-2023-05298). J.M.A. gratefully acknowledge the support from INCT Materials Informatics (CNPq 406447/2022-5).

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## Data Availability Statement

The data supporting this article have been included as part of the Supplementary Information. This file contains all original data for the DFT calculations presented in the main manuscript Figure 8.