



Cite this: *Chem. Commun.*, 2024, 60, 9254

Received 10th June 2024,  
Accepted 26th July 2024

DOI: 10.1039/d4cc02609h

rsc.li/chemcomm

# Versatile dehydrogenation of carbonyls enabled by an iodine(III) reagent†

Bence B. Botlik, Patrick Finkelstein, Ann-Sophie K. Paschke, Julia C. Reisenbauer and Bill Morandi\*

**We report the utilisation of an iodine(III) reagent to access  $\alpha,\beta$ -unsaturated carbonyls from the corresponding silyl enol ethers of ketones and aldehydes, and from enol phosphates of lactones and lactams. The transformation is rapid, scalable, and can be carried out in one pot, directly dehydrogenating saturated carbonyls.**

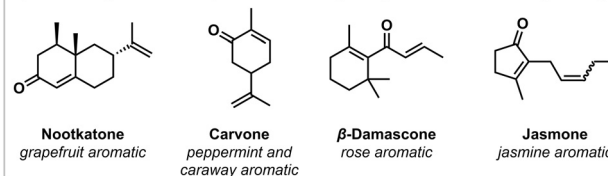
Converting saturated carbonyl compounds into the corresponding  $\alpha,\beta$ -unsaturated derivatives is a key transformation in synthetic organic chemistry.<sup>1</sup> The products of these dehydrogenation processes are valuable intermediates and versatile building blocks in the synthesis of natural products as well as industrially relevant targets, and ubiquitous functional moieties in various materials.<sup>2,3</sup> The relevance of  $\alpha,\beta$ -unsaturated carbonyls is particularly reflected by their abundance among aroma compounds and odorants, thus they are of special interest for the fragrance industry (Scheme 1A).<sup>4</sup>

In addition, their important role as useful synthetic intermediates is showcased by the large number of established interconversions for these functional handles that further enable their derivatisation. For instance, conjugate additions, cross couplings, regioselective  $\alpha$ -functionalisations, and alkene difunctionalisation reactions open different avenues to access a broad range of products from  $\alpha,\beta$ -unsaturated carbonyls.<sup>2,5</sup>

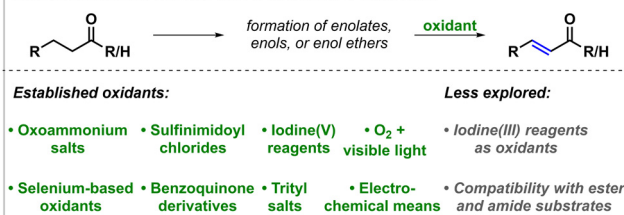
Most known dehydrogenation methods formally proceed *via* two consecutive chemical transformations. Typically, in the first step, an enol or an enolate is formed, which is either directly transformed into the desired unsaturated product, or is trapped *in situ* generating the corresponding silyl enol ether. This strategy allows for the activation of the  $\alpha$ -position and controls the regioselectivity of the subsequent dehydrogenation step mediated by an oxidant, resulting in the formation of the corresponding desaturated products. To date, most of the developed carbonyl  $\alpha,\beta$ -dehydrogenation reactions employ transition metals.<sup>1,6–15</sup> In

contrast, transition metal-free approaches are comparatively less established. Besides a recent electrochemical approach,<sup>16</sup> other common strategies rely on the utilisation of oxoammonium salts,<sup>17</sup> benzoquinone derivatives,<sup>18</sup> selenium compounds,<sup>19</sup> trityl salts<sup>20</sup> or sulfinimidoyl derivatives<sup>21</sup> as oxidants (Scheme 1B). Additionally, methods implementing iodine(V) reagents have also been reported, both using either silyl enol ethers or the saturated carbonyls as starting scaffolds.<sup>22,23</sup> However, these methods are often limited in their scopes to the dehydrogenation of ketones and aldehydes, and commonly require long reaction times and elevated temperatures. Methods relying on iodine(III) reagents as

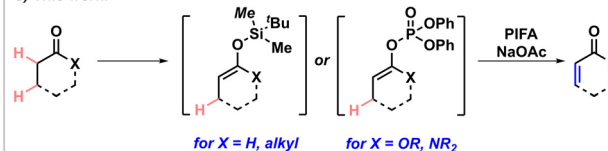
## A) Industrially relevant compounds containing an $\alpha,\beta$ -unsaturated carbonyl moiety:



## B) Transition metal-free $\alpha,\beta$ -dehydrogenation of carbonyls:



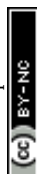
## C) This work:



**Scheme 1** (A) Relevance of  $\alpha,\beta$ -unsaturated carbonyl compounds. (B) Previous examples of transition metal-free carbonyl dehydrogenation. (C) This work.

Laboratorium für Organische Chemie, ETH Zürich, Vladimir Prelog Weg 3, HCI, 8093 Zürich, Switzerland. E-mail: morandib@ethz.ch

† Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d4cc02609h>



oxidants are scarce, and to date have been restricted to specific systems, such as the flavanone to flavone transformations,<sup>24</sup> and the double dehydrogenation of cyclic  $\beta$ -ketocarboxyls.<sup>25</sup> Moreover, silyl enol ethers usually exhibit different reactivities with iodine(III) reagents and in most cases result in the formation of  $\alpha$ -functionalised ketones, commonly generating  $\alpha$ -ketoacetate derivatives,<sup>26</sup>  $\alpha$ -hydroxyketones,<sup>27</sup>  $\alpha$ -methoxyketones,<sup>28</sup> and  $\alpha$ -sulfonylketones.<sup>29</sup> Our group recently disclosed that iodine(III) reagents facilitate the nitrogen atom insertion into the silyl enol ether of indanone and related scaffolds in the presence of an external nitrogen atom source.<sup>30–33</sup> Based on these results, similar substrates possessing silyl enol ether functional handles were investigated towards their reactivity with iodine(III) reagents. However, a mechanistically distinct reactivity was observed in various cases, leading to the formation of the corresponding dehydrogenation products. This is hypothesised to be the result of utilising the understudied combination of a particularly strong hypervalent iodine oxidant and an equivalent amount of base, with easily oxidisable silyl enol ether reaction partners. Herein, the optimisation and development of a general, rapid, and operationally simple iodine(III)-mediated dehydrogenation method is presented, which grants access to  $\alpha,\beta$ -unsaturated carbonyls from silyl enol ethers of ketones and aldehydes, and from enol phosphates of lactones and lactams (Scheme 1C).

Initial optimisation of the transformation was performed using *tert*-butyldimethylsilyl-(TBS)-protected cyclohexenol **1a** as the model substrate (Table 1). Acetone was found to be the most effective solvent for the reaction, with acetonitrile and tetrahydrofuran (THF) also providing moderate yields of the dehydrogenation product **2a**, while other examined solvents showed significantly lower conversion to the desired product. *Bis*(trifluoroacetoxy)iodobenzene (PIFA) was the most suitable oxidant, as the use of other, less oxidising iodine(III) reagents either afforded the desired product in low yields or no product

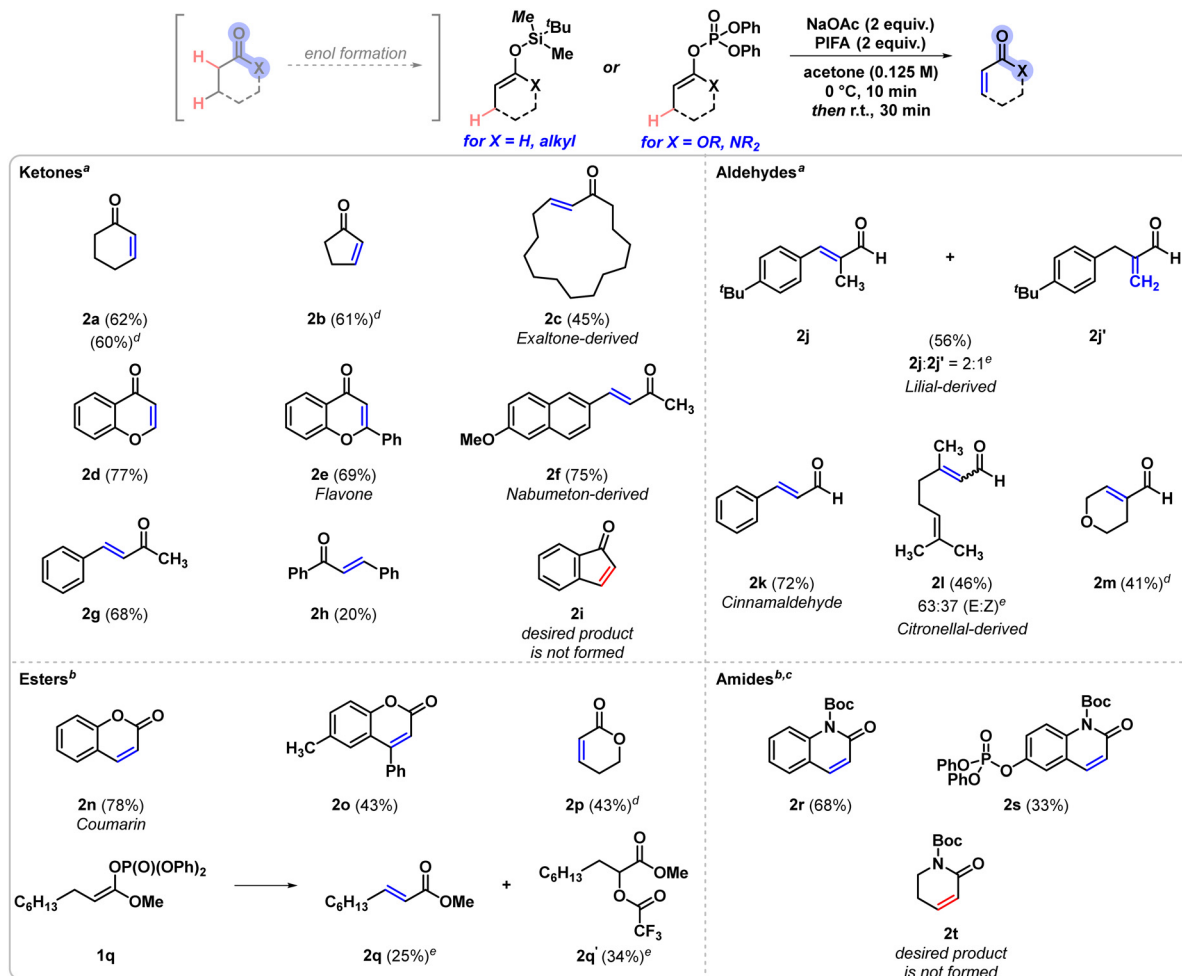
formation was detected. In addition, various inorganic bases were found to be compatible with the initial screening conditions, resulting in nearly identical yields, however, amine bases reduced the desired reactivity. Using oxidant loadings lower than 2 equivalents was detrimental to the yield, however, further increasing the amount of oxidant in the reaction also did not improve the overall yield. Variation of the silyl groups incorporated in the starting material indicated the superiority of the TBS group, as *tri*-isopropylsilyl (TIPS) enol ethers were transformed with significantly lower yields, and in general low reactivity was observed in the case of other silyl groups, such as *tri*-methylsilyl and *tri*-ethylsilyl groups. Finally, the effects of molarity and temperature were examined, and the highest yields of product **2a** were obtained when the reaction was performed at 0 °C using a more dilute solution (0.125 M with respect to substrate **1a** in acetone). With the optimal conditions in hand, we examined the scope of the transformation (Scheme 2). Cyclic aliphatic ketones **2a** and **2b** were both obtained in approximately 60% <sup>1</sup>H NMR yield, which corresponds well to the isolated yield of **2a**. Similarly, other silyl enol ethers exhibiting larger ring sizes were also found to be compatible with the reaction and the macrocyclic product **2c** was isolated in moderate 45% yield. Chromanone derivatives **2d** and **2e** were obtained in good yields (77% and 69%, respectively), as well as phenyl-conjugated alkenes, including **2g** and Nabumeton-derived product **2f**. In contrast, chalcone **2h** was isolated in lower yield, likely due to the decomposition of the silyl enol ether under the reaction conditions, resulting the saturated carbonyl as a major side product. In the case of indanone **1i**, however, the corresponding dehydrogenation product **2i** could not be observed. Aldehyde-derived starting scaffolds exhibited comparable reactivity to that of ketone-derived ones. The compatibility of aldehydes with our method was examined through a number of examples that are important to the fragrance industry.<sup>4</sup> Linal-derived **2j** and **2j'** were obtained in a 2 : 1 ratio of regioisomers and were isolated in a combined 56% yield, whereas cinnamaldehyde **2k** was isolated in 72% yield. Similarly, citronellal **1l** transformed into dehydrogenated product **2l** in the presence of another alkene handle in moderate yield, resulting in the formation of both the *E* and *Z* alkenes in a 2 : 1 ratio. Finally, the formation of the desaturated aliphatic aldehyde **2m** was also observed. Apart from ketone and aldehyde substrates, lactones and lactams were also found to be suitable substrates for the reaction. As the silyl enol ethers of lactones and lactams are unstable and prone to hydrolysis, the more stable enol phosphate derivatives were employed. Coumarin (**2n**) was obtained in high yield (78%), while its substituted derivative **2o** was isolated in reduced yield. The desaturation of a non-conjugated aliphatic lactone could also be achieved, resulting in the formation of product **2p** in 43% <sup>1</sup>H NMR yield. When employing an open chain enol phosphate, the desired product **2q** was formed and co-isolated with the  $\alpha$ -trifluoroacetylated side product **2q'**. Additional examples showcasing the reaction's translation to enol phosphates derived from lactams were tested and the corresponding dehydrogenation products **2r** and **2s** were isolated in

**Table 1** Selected optimisation data for the dehydrogenation of silyl enol ethers via iodine(III) reagents

| Entry | Deviation from above <sup>a</sup>  | Yield <sup>b</sup> of <b>2a</b> (%) |
|-------|------------------------------------|-------------------------------------|
| 1     | None                               | 60                                  |
| 2     | MeCN instead of acetone            | 38                                  |
| 3     | THF instead of acetone             | 43                                  |
| 4     | PIDA instead of PIFA               | 0                                   |
| 5     | HTIB instead of PIFA               | 5                                   |
| 6     | KOH instead of NaOAc               | 59                                  |
| 7     | KTFA instead of NaOAc              | 58                                  |
| 8     | Et <sub>3</sub> N instead of NaOAc | 48                                  |
| 9     | 1 equiv. PIFA                      | 42                                  |
| 10    | 3 equiv. PIFA                      | 60                                  |
| 11    | TIPS instead of TBS                | 28                                  |

<sup>a</sup> Reaction conditions: silyl enol ether (0.05 mmol), PIFA (0.10 mmol), NaOAc (0.10 mmol), acetone-*d*<sub>6</sub> (0.125 M), 0 °C, 10 min, then room temperature (r.t.), 30 min. <sup>b</sup> Yields determined by <sup>1</sup>H NMR analysis of the crude reaction mixtures, using mesitylene as the internal standard.





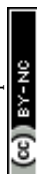
**Scheme 2** Substrate scope. The yields are isolated yields of the reaction, unless indicated otherwise. <sup>a</sup>Silyl enol ethers were used as starting materials. <sup>b</sup>Enol phosphates were used as starting materials. <sup>c</sup>The reaction was carried out in acetonitrile instead of acetone. <sup>d</sup>Reaction carried out on 0.05-mmol scale, yields determined by <sup>1</sup>H NMR analysis of the crude reaction mixtures, using mesitylene internal standard. <sup>e</sup>Isomeric ratios were determined by <sup>1</sup>H NMR analysis. For details, see ESI.†

68% and 33%, respectively. The non-conjugated enol phosphate **1t**, however, did not yield the corresponding dehydrogenation product **2t**.

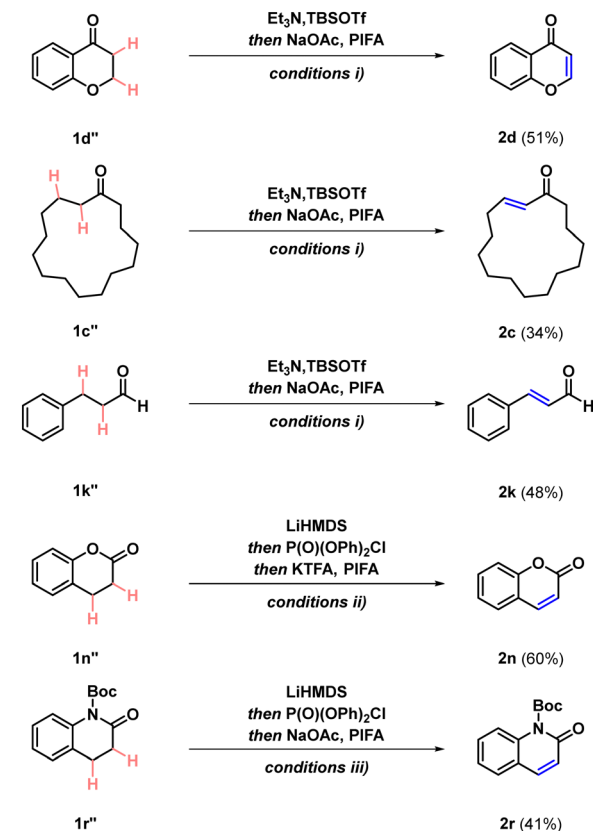
We hypothesised that the reaction could be carried out in a one-pot manner directly using the saturated carbonyl compounds as starting materials, which would greatly improve the synthetic utility of the developed transformation. Therefore, we set out to examine the one-pot dehydrogenation through a representative example of each four carbonyl classes (Scheme 3A). To our delight, ketone **1d''** yielded the desired product **2d** in 51% yield upon using THF solvent for the silyl enol ether formation step and a THF:acetone = 1:9 solvent mixture in the subsequent dehydrogenation step. Similarly, the direct dehydrogenation of carbocyclic ketone **1c''** was also demonstrated. Using the same conditions, aldehyde **1k''** afforded the corresponding dehydrogenation product **2k** in 48% isolated yield. Furthermore, the possibility for the one-pot dehydrogenation of lactones was demonstrated by using **1n''**, which afforded product **2n** in 60% isolated yield. Key to success was the implementation of THF as the reaction solvent,

as it is compatible with both the enol phosphate formation and the subsequent oxidation step, alongside using increased amounts of PIFA and implementing KTFA as the base. By using an analogous strategy, the dehydrogenation of lactam **1r''** also proceeded in a one-pot fashion to afford dehydrogenated lactam **2r**. These results possess considerable synthetic importance, as transition metal-free dehydrogenation methods of lactones and lactams are scarce.<sup>1,34</sup> Even though higher overall yields are observed when using the two-step approach, the combined yields over the two steps are comparable to the yields of the one-pot transformation. The scalability of the reaction was also examined, and the formation of cinnamaldehyde (**2k**) was conducted on a gram-scale (14.0 mmol of starting material; Scheme 3B), providing the desired product in 73% isolated yield, which is identical to the result obtained for the 1-mmol scale.

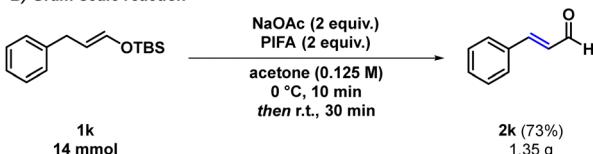
In conclusion, we have developed an operationally simple carbonyl dehydrogenation method, complementing already existing methodologies for the formation of  $\alpha,\beta$ -unsaturated ketones, aldehydes, lactones, and lactams, by using a readily available iodine(III) oxidant under basic conditions. Furthermore, we have



## A) One-pot dehydrogenation reactions



## B) Gram-scale reaction



**Scheme 3** Further experiments. (A) One-pot dehydrogenation reactions. Conditions: (i)  $\text{Et}_3\text{N}$  (1.5 equiv.), TBSOTf (1.2 equiv.), THF (1 M), r.t., 1 h; then NaOAc (2 equiv.), PIFA (2 equiv.), THF:acetone = 1 : 9 (0.1 M), 0 °C, 10 min; then r.t., 30 min. (ii) LiHMDS (1.0 equiv.), dry THF (0.125 M), –78 °C, 30 min; then  $\text{P(O)(OPh)}_2\text{Cl}$  (1.0 equiv.), –78 °C to r.t., 1 h; then KTFA (4 equiv.), PIFA (4 equiv.), r.t., 30 min. (iii) LiHMDS (1.1 equiv.), dry THF (0.125 M), –78 °C, 30 min; then  $\text{P(O)(OPh)}_2\text{Cl}$  (1.1 equiv.), –78 °C to r.t., 1 h; then NaOAc (2 equiv.), PIFA (2 equiv.), THF:MeCN = 1 : 1 (0.06 M), 0 °C, 10 min; then r.t., 30 min. (B) Scale-up experiment.

demonstrated the excellent scalability of the reaction to gram-scale as well as the feasibility of a one-pot approach to rapidly access  $\alpha,\beta$ -desaturated products directly from the corresponding saturated carbonyls.

We thank Prof. Ori Green (Technion) for valuable discussions. This work was supported by ETH Zürich and the Swiss National Science Foundation (SNSF 184658). B. B. B. and J. C. R. acknowledge a fellowship from the Scholarship Fund of the Swiss Chemical Industry (SSCI). A. S. P. acknowledges a fellowship from the Fonds der Chemischen Industrie (FCI). We thank the NMR and MS (MoBIAS) service departments at ETH Zürich for technical assistance and the Morandi group for critical proofreading of the manuscript.

## Data availability

The data supporting this article have been included as part of the ESI.†

## Conflicts of interest

There are no conflicts to declare.

## Notes and references

- S. Gnaïm, J. C. Vantourout, F. Serpier, P.-G. Echeverria and P. S. Baran, *ACS Catal.*, 2021, **11**, 883–892.
- H. Chen, L. Liu, T. Huang, J. Chen and T. Chen, *Adv. Synth. Catal.*, 2020, **362**, 3332–3346.
- G. P. Rosa, A. M. L. Seca, M. do, C. Barreto and D. C. G. A. Pinto, *ACS Sustainable Chem. Eng.*, 2017, **5**, 7467–7480.
- D. Pybus and C. Sell, *The Chemistry of Fragrances*, Royal Society of Chemistry, 2007.
- S. Patai, *The Chemistry of enones*, Wiley, 1989.
- Y. Chen, J. P. Romaine and T. R. Newhouse, *J. Am. Chem. Soc.*, 2015, **137**, 5875–5878.
- D. Huang, Y. Zhao and T. R. Newhouse, *Org. Lett.*, 2018, **20**, 684–687.
- M. Chen and G. Dong, *J. Am. Chem. Soc.*, 2017, **139**, 7757–7760.
- M.-M. Wang, X.-S. Ning, J.-P. Qu and Y.-B. Kang, *ACS Catal.*, 2017, **7**, 4000–4003.
- Y. Shang, X. Jie, K. Jonnada, S. N. Zafar and W. Su, *Nat. Commun.*, 2017, **8**, 2273.
- M. Chen, A. J. Rago and G. Dong, *Angew. Chem., Int. Ed.*, 2018, **57**, 16205–16209.
- Z. Wang, Z. He, L. Zhang and Y. Huang, *J. Am. Chem. Soc.*, 2018, **140**, 735–740.
- D. Huang, S. M. Szweczyk, P. Zhang and T. R. Newhouse, *J. Am. Chem. Soc.*, 2019, **141**, 5669–5674.
- Y. Zhao, Y. Chen and T. R. Newhouse, *Angew. Chem., Int. Ed.*, 2017, **56**, 13122–13125.
- D. Huang, S. M. Szweczyk, P. Zhang and T. R. Newhouse, *J. Am. Chem. Soc.*, 2019, **141**, 5669–5674.
- S. Gnaïm, Y. Takahira, H. R. Wilke, Z. Yao, J. Li, D. Delbrayelle, P.-G. Echeverria, J. C. Vantourout and P. S. Baran, *Nat. Chem.*, 2021, **13**, 367–372.
- M. Hayashi, M. Shibuya and Y. Iwabuchi, *Org. Lett.*, 2012, **14**, 154–157.
- E. A. Braude, L. M. Jackman and R. P. Linstead, *J. Chem. Soc.*, 1954, 3548–3563.
- K. B. Sharpless, R. F. Lauer and A. Y. Teranishi, *J. Am. Chem. Soc.*, 1973, **95**, 6137–6139.
- M. E. Jung, Y.-G. Pan, M. W. Rathke, D. F. Sullivan and R. P. Woodbury, *J. Org. Chem.*, 1977, **42**, 3961–3963.
- T. Mukaiyama, J. Matsuo and H. Kitagawa, *Chem. Lett.*, 2000, 1250–1251.
- K. C. Nicolaou, D. L. F. Gray, T. Montagnon and S. T. Harrison, *Angew. Chem., Int. Ed.*, 2002, **41**, 996–1000.
- K. C. Nicolaou, Y.-L. Zhong and P. S. Baran, *J. Am. Chem. Soc.*, 2000, **122**, 7596–7597.
- O. Prakash, S. Pahuja and R. M. Moriarty, *Synth. Commun.*, 1990, **20**, 1417–1422.
- S.-S. Liu, L. Wang, Y.-N. Duan, A. Yu and C. Zhang, *Sci. China: Chem.*, 2019, **62**, 597–601.
- I. Brunolevskaya, K. Kusainova and A. Kashin, *Zh. Org. Khim.*, 1988, **24**, 358.
- R. M. Moriarty, M. P. Duncan and O. Prakash, *J. Chem. Soc., Perkin Trans. 1*, 1987, 1781–1784.
- R. M. Moriarty, O. Prakash, M. P. Duncan, R. K. Vaid and H. A. Musallam, *J. Org. Chem.*, 1987, **52**, 150–153.
- R. M. Moriarty, R. Penmasta, A. K. Awasthi, W. R. Epa and I. Prakash, *J. Org. Chem.*, 1989, **54**, 1101–1104.
- P. Finkelstein, J. C. Reisenbauer, B. B. Botlik, O. Green, A. Florin and B. Morandi, *Chem. Sci.*, 2023, **14**, 2954–2959.
- J. C. Reisenbauer, O. Green, A. Franchino, P. Finkelstein and B. Morandi, *Science*, 2022, **377**, 1104–1109.
- J. C. Reisenbauer, A.-S. K. Paschke, J. Krizic, B. B. Botlik, P. Finkelstein and B. Morandi, *Org. Lett.*, 2023, **25**, 8419–8423.
- B. B. Botlik, M. Weber, F. Ruepp, K. Kawanaka, P. Finkelstein and B. Morandi, *Angew. Chem., Int. Ed.*, 2024, e202408230.
- P. Spieß, M. Berger, D. Kaiser and N. Maulide, *J. Am. Chem. Soc.*, 2021, **143**, 10524–10529.

