Environmental Science: Atmospheres

ROYAL SOCIETY OF CHEMISTRY

PAPER

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Cite this: *Environ. Sci.: Atmos.*, 2023, **3**, 1585

The driving effects of common atmospheric molecules for formation of clusters: the case of sulfuric acid, nitric acid, hydrochloric acid, ammonia, and dimethylamine†

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Understanding how secondary aerosols form in the atmosphere is one of the main uncertainties for a better understanding of global warming. Secondary aerosols form from gas-phase molecules that combine to create prenucleation complexes, which can then grow to form aerosols. The study of the formation of prenucleation complexes is difficult from both an experimental and theoretical point of view. Sulfuric acid has been linked to the formation of aerosols, yet the details of interactions are not understood. We have completed an exhaustive study of the formation of prenucleation complexes of three strong acids: sulfuric acid, nitric acid, and hydrochloric acid, combined with ammonia and dimethylamine bases, and three water molecules. By combining an evolutionary algorithm search routine with density functional geometry optimizations and single-point electronic energy calculations with complete basis set (CBS) extrapolations, we have completed an exhaustive search of the DLPNO-CCSD(T)/CBS//ωB97X-D/6-31++G** Gibbs free energy surface for this system. We have used previous work where the weaker formic acid replaces either nitric acid or hydrochloric acid to explore the details of how three acids combine with two bases and a few water molecules to make prenucleation clusters. As clusters grow, stabilizing effects of nitric acid, hydrochloric acid, and formic acid change in unique ways. This research adds to the body of work that illustrates that, depending on the system being studied, the acid/base strength of the monomers, the charge distribution within the clusters, and the detailed hydrogen bond topology have a subtle interplay that determines which cluster is most stable.

Received 28th July 2023
Accepted 3rd September 2023

DOI: 10.1039/d3ea00118k

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Environmental significance

The impact of aerosols on global climate is a major source of uncertainty in the understanding of global warming. Understanding the growth and formation of secondary aerosols is key to fully grasp the effect of aerosols on Earth's climate. The beginning stages in the formation of prenucleation complexes, that eventually lead to larger aerosols, cannot currently be investigated experimentally. Using robust computational chemistry methods, we have determined the lowest Gibbs free energy clusters and predicted equilibrium concentrations of the sulfuric acid–nitric acid–hydrochloric acid–ammonia–dimethylamine–water system. When compared to previously studied systems sulfuric acid–formic acid–hydrochloric acid–ammonia–dimethylamine–water and sulfuric acid–formic acid–nitric acid–ammonia–dimethylamine–water, we see that the stability of the nitric acid–hydrochloric acid is affected by the number and identity of surrounding monomers.

1 Introduction

Aerosols' impact on Earth's climate is one of the main sources of uncertainty for a better understanding of global warming.[1] They affect the climate by scattering light, absorbing and emitting radiation, and as cloud condensation nuclei (CCN).[2] CCN nucleate cloud droplets and ice particles, and thus dramatically influence the processes that lead to the formation of rain, snow, hail, and other forms of precipitation.[3] Aerosols exert a net cooling of the atmosphere, but the uncertainties in our knowledge are larger than the cooling effects.[1–4] Primary aerosols such as sea spray and dust enter the atmosphere directly and secondary aerosols form in the gas phase from precursor molecules.[5] Secondary aerosols are estimated to account for 50% or more of all aerosol particles, and perhaps even all CCN-candidate particles.[6,7] However, how secondary aerosols form is still not well understood, as experimental and

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† Electronic supplementary information (ESI) available: Figures of minimum energy structures, sequential hydration tables, DFT and DLPNO-CCSD(T) energy tables, optimized structures, derivation of CBS extrapolation formula, and example calculations of stepwise energies. See DOI: https://doi.org/10.1039/d3ea00118k

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theoretical approaches for learning how gaseous vapor from multiple molecular species combine to create the initial pre-nucleation clusters is a very difficult problem.[8–14] Thus, an incredible amount of experimental and computational effort has been devoted to understanding which species are involved, and how they combine together, to form secondary aerosols.[6,8–85] It is common in the field to work under the assumption that acids and bases grow into an ensemble that contains an equal number of both types of molecules.[8–14] Yet, the tremendous number of potential compounds that can form prenucleation clusters is vast, and there is no experimental technique that can identify all of the different molecules in prenucleation complexes or small aerosols in the sub-nanometer (nm) to nm size range.[12] Thus, this problem has been described as seeking a needle in a haystack.[14] Our recent examination of prenucleation complexes which contain formic acid, a weaker acid compared to sulfuric acid, nitric acid, and hydrochloric acid, in clusters of unequal amounts of acids and bases, reveal that the detailed hydrogen bonding topology of a particular cluster is sometimes more important than traditional acid/base strength.[65,67,72,86] In our previous work we have simulated how sulfuric acid (SA), formic acid (FA), ammonia (A), and water (W), form prenucleation complexes and discovered that FA is as effective as A at forming clusters with SA.[65] We have also computed all the interactions between SA, A, and the amino acids glycine and serine, and found that the amino acids and A are practically interchangeable, and there is no easy way to predict which of the three molecules are protonated when SA loses its proton to one of those species. The acid/base reaction proceeds with all three species at specific instances that depend on the details of the molecular cluster.[67] We next explored the role of three acids, two bases, and a handful of water molecules to form prenucleation complexes by computing all possible combinations of one SA, one FA, one nitric acid (NA), one A, and one dimethylamine (DMA) with five waters (W). We again found that the detailed geometries of each Gibbs free energy minimum is often more important than traditional acid or base strength.[72] Addition of water to a dry cluster can enhance stabilization, as the (SA)(NA)(A)(DMA)(W) cluster has enhanced stability. NA is predicted to drive prenucleation almost as efficiently as SA.[72] Most recently, we swapped NA for hydrochloric acid (HCl), and explored all possible combinations of SA, FA, HCl, A, DMA, and up to three W to probe the effects of this swap of one strong acid for another.[86] This first detailed study of HCl interacting with two other acids and two bases revealed that NA forms stronger interactions in dry clusters than HCl, but sometimes as the clusters grow larger and are hydrated the sequential Gibbs free energies of clusters containing HCl become more favorable than those for NA.[86] Overall this previous work adds to the body of evidence that reveals that hydrogen bond topology and the detailed structural interactions within a prenucleation cluster are as important as conventional ideas such as acid/base strength. Prenucleation clusters, held together by hydrogen bonding and other van der Waals forces, have free energies determined by the subtle interactions between enthalpy and entropy, which depend on the molecular structure of the complex. In this paper, we present the results for swapping FA with NA. This detailed exploration of the Gibbs free energy surface of (SA)(NA)(HCl)(A)(DMA)$(W)_{0-3}$ allows us to compare all three systems, such that we have examined each combination of these three acids with these two bases where SA is one of the acids.

2 Methodology

Configurational sampling of clusters containing SA, NA, HCl, A, DMA, and up to three W was employed using the genetic-algorithm-based OGOLEM program,[87,88] followed by semi-empirical geometry optimizations with the GFN2-xTB semi-empirical method.[89–91] An evolutionary algorithm in OGOLEM rearranges the initial configurations until reaching convergence on a final set of clusters in the designated pool. We used a pool size of 1000 and set the number of global GFN2-xTB optimizations to 100 000. Because of the uncertainties associated with semi-empirical methods,[58,92] all structures within 30 kcal mol^{-1} of the GFN2-xTB global minimum were then re-optimized at the ωB97X-D/6-31++G** level of theory[93–98] using the Gaussian 16 Rev. B01 program,[99] and duplicate structures were removed.[100] All DFT structures within 8 kcal mol^{-1} of the ωB97X-D/6-31++G** electronic energy global minimum were recalculated using the domain-based local pair natural orbital coupled cluster (DLPNO-CCSD(T)) method[101–108] with single, double, and semi-canonical perturbative triple excitations with three Dunning basis sets,[109–112] cc-pVnZ (n = D, T, Q) using Orca 5.0.1.[113] Estimates of the thermodynamic corrections for $H°$, $S°$, and $G°$ were computed at a standard state of 1 atm pressure and temperatures of 216.65, 273.15, and 298.15 K using the THERMO.pl script[114] from the National Institute of Science and Technology. These computations used the ωB97X-D/6-31++G** frequencies, which were scaled by a factor of 0.971 to partially account for anharmonicity.[115,116] An inverse 4–5 polynomial complete basis set (CBS) extrapolation[117] on the three CCSD(T)/cc-pVnZ (n = D, T, Q) electronic energies was used to obtain CBS electronic energies. The CBS energies were combined with the ωB97X-D/6-31++G** thermal corrections to obtain the final $G°$ values. The complete methodology is outlined in Fig. 1. Once the final $G°$ values were calculated for the individual clusters, $\Delta G°_{binding}$ was calculated as follows:

$$\Delta G^{\circ}_{blind} = G^{\circ}_{cluster} - \sum G^{\circ}_{monomers} \quad (1)$$

The sequential binding energies for the addition of a monomer were computed using:

$$\Delta G^{\circ}_{addition} = G^{\circ}_{new\ cluster} - \left(\Delta G^{\circ}_{blind} + G^{\circ}_{new\ monomer}\right) \quad (2)$$

These addition energies, alongside initial concentrations for the monomers, were used to calculate the equilibrium concentrations of each cluster, assuming a closed system. Using a closed system accounts for monomer vapor depletion. These concentrations were calculated at 216.65 K and 298.15 K, which correspond to the top and bottom of the troposphere. We used a water concentration of 7.7×10^{17} cm^{-3} at 298 K and 9.9×10^{14}

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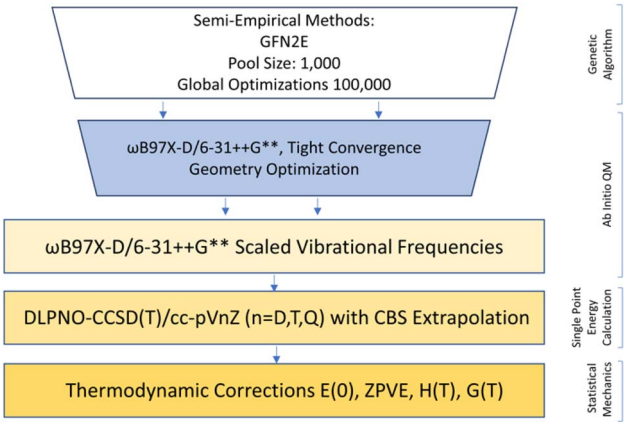

Fig. 1 Complete methodology used in this study.

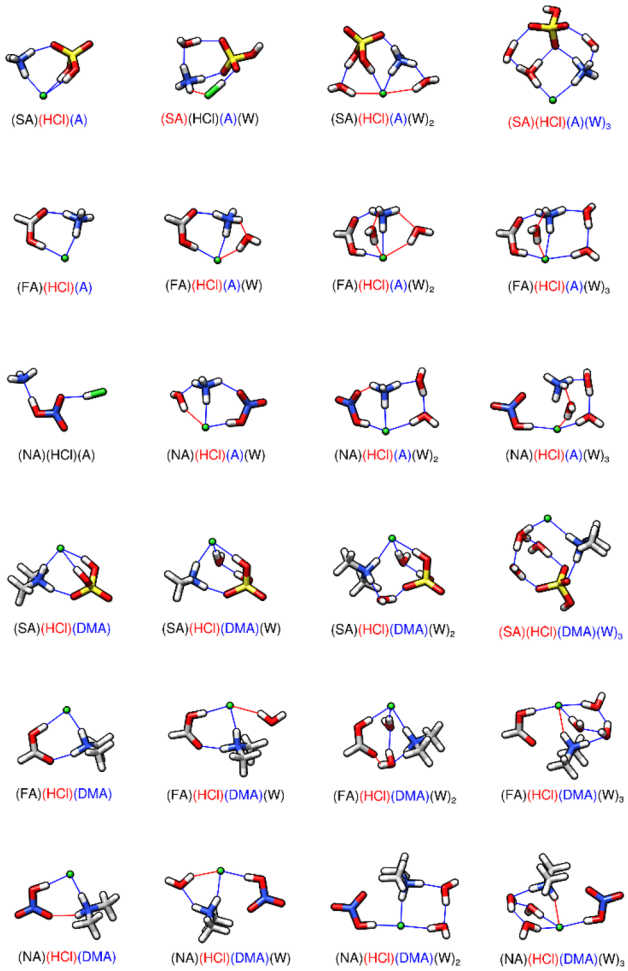

Fig. 2 DLPNO-CCSD(T)/CBS//ωB97X-D/6-31++G** minimum energy structures for the sequential hydration of hydrochloric acid with one acid and one base at 298 K. The molecule labels are colored according to charge as follows: blue = +1, black = 0, red = −1. Atoms are drawn in the following colors: hydrogen – white, carbon – grey, nitrogen – blue, oxygen – red, sulfur – yellow, chlorine – green. In the figures, hydrogen bonds are marked in blue, which have hydrogen-bonded distances of less than 2.2 Å and hydrogen bond angles between 140 and 180°. Red lines are used to denote van der Waals forces where either the bond angle encompassing the hydrogen is less than 140° or the hydrogen-bond distance is greater than 2.2 Å.

cm^{-3} at 217 K, which corresponds to 100% humidity at the bottom and top of the troposphere.[5] Initial starting concentrations of the monomers were 5×10^7 cm^{-3} for SA, 9.8×10^{10} cm^{-3} for NA, 1×10^9 cm^{-3} for HCl, 2×10^{11} cm^{-3} for A, and 2×10^9 cm^{-3} for DMA at 298 K, which are relevant for each of the monomers over inland and urban areas.[5,30,118–125] At 217 K, the monomer concentrations are decreased by three orders of magnitude to compensate for the reduction of condensable/nucleating/aerosol-forming vapors in the upper troposphere. This is a rough estimate based on the three orders of magnitude decrease in concentration of water, and we used this approximation since experimental concentrations of these monomers in the upper troposphere are difficult to measure. These results were combined with results from the previously studied $(SA)(FA)(NA)(A)(DMA)(W)_{0-3}$ and $(SA)(FA)(HCl)(A)(DMA)(W)_{0-3}$ clusters,[72,86] allowing for a comparison of the different combinations of three different acids with the two different bases. We note that there are several limitations to our methodology.[14] First, as we are unable to know if we have sampled the entirety of the semi-empirical potential energy surface (PES), there is always a possibility that minimum energy clusters were missed. Second, computing the CCSD(T) electronic energies on the DFT structures means we are making electron correlation corrections to the original DFT PES. Third, using the DLPNO approximation, and scaled frequencies from the rigid-rotor harmonic oscillator model generally leads to slightly more positive free energies. Finally, we note that we have previously investigated the effect of augmented bases sets (aug-cc-pVnZ) for the SA–A–glycine–serine system and found that the CBS extrapolation converged both results to within a half-kcal mol^{-1}.[67] More work in this area is warranted. For more details about our methodology, we refer the reader to our recent review.[14]

3 Results and discussion

3.1 Trimers of two acids and one base

For the $(SA)(NA)(HCl)(A)(DMA)(W)_{0-3}$ system, all monomers, dimers, and one acid-two base trimers have previously been discussed and are listed in the ESI (Tables T1–T5 and Fig. S1–S5†).[72,86] We have not included the two acid dimers containing HCl, $[(SA)(HCl)(W)_{0-3}$, $(NA)(HCl)(W)_{0-3}$, and $(FA)(HCl)(W)_{0-3}]$ as previous results have shown these clusters are not atmospherically significant due to the lack of hydrogen bonding sites and bases present for deprotonation.[72,86] The (SA)(NA)(HCl) trimer structure is included in the ESI† along with the energetics for the hydration of all the three acid clusters (Fig. S6 and Table T6†). We begin the discussion of the $(SA)(NA)(HCl)(A)(DMA)(W)_{0-3}$ system with two acid-one base trimers. The structures and energies of the two acid-one base trimers are shown in Fig. 2 and Tables 1 & 2. In the figures, hydrogen bonds are marked in blue, and are defined as hydrogen-bonded distances of less than 2.2 Å and hydrogen bond angles between 140 and 180°. Red lines are used to denote van der

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Table 1 DLPNO-CCSD(T)/CBS//ωB97X-D/6-31++G** Gibbs free energy changes (kcal mol^{-1}) associated with the formation and sequential hydration of two acid and ammonia trimers at atmospherically relevant temperatures and 1 atm pressure. Values from a[72] and b[86]

Cluster	216.65 K	273.15 K	298.15 K
SA + FA + A ⇌ (SA)(FA)(A)[a]	−17.13	−13.39	−11.74
(SA)(FA)(A) + W ⇌ (SA)(FA)(A)(W)[a]	−4.65	−2.76	−1.93
(SA)(FA)(A)(W) + W ⇌ (SA)(FA)(A)$(W)_2$[a]	−3.19	−1.36	−0.55
(SA)(FA)(A)$(W)_2$ + W ⇌ (SA)(FA)(A)$(W)_3$[a]	−3.23	−1.42	−0.62
SA + NA + A ⇌ (SA)(NA)(A)[a]	−15.74	−11.99	−10.40
(SA)(NA)(A) + W ⇌ (SA)(NA)(A)(W)[a]	−5.73	−4.04	−3.23
(SA)(NA)(A)(W) + W ⇌ (SA)(NA)(A)$(W)_2$[a]	−3.83	−2.13	−1.38
(SA)(NA)(A)$(W)_2$ + W ⇌ (SA)(NA)(A)$(W)_3$[a]	−3.15	−1.51	−0.79
SA + HCl + A ⇌ (SA)(HCl)(A)[b]	−13.15	−9.52	−7.90
(SA)(HCl)(A) + W ⇌ (SA)(HCl)(A)(W)[b]	−3.92	−2.50	−1.88
(SA)(HCl)(A)(W) + W ⇌ (SA)(HCl)(A)$(W)_2$[b]	−3.54	−1.39	−0.45
(SA)(HCl)(A)$(W)_2$ + W ⇌ (SA)(HCl)(A)$(W)_3$[b]	−3.69	−2.06	−1.33
FA + NA + A ⇌ (FA)(NA)(A)[a]	−8.42	−5.25	−3.87
(FA)(NA)(A) + W ⇌ (FA)(NA)(A)(W)[a]	−2.79	−0.32	0.78
(FA)(NA)(A)(W) + W ⇌ (FA)(NA)(A)$(W)_2$[a]	−2.84	−0.96	−0.18
(FA)(NA)(A)$(W)_2$ + W ⇌ (FA)(NA)(A)$(W)_3$[a]	−0.80	0.69	1.40
FA + HCl + A ⇌ (FA)(HCl)(A)[b]	−6.65	−3.04	−1.45
(FA)(HCl)(A) + W ⇌ (FA)(HCl)(A)(W)[b]	−4.41	−2.80	−2.10
(FA)(HCl)(A)(W) + W ⇌ (FA)(HCl)(A)$(W)_2$[b]	−3.53	−1.82	−1.17
(FA)(HCl)(A)$(W)_2$ + W ⇌ (FA)(HCl)(A)$(W)_3$[b]	−3.31	−1.70	−0.88
NA + HCl + A ⇌ (NA)(HCl)(A)	−4.73	−1.99	−0.78
(NA)(HCl)(A) + W ⇌ (NA)(HCl)(A)(W)	−5.30	−3.06	−2.07
(NA)(HCl)(A)(W) + W ⇌ (NA)(HCl)(A)$(W)_2$	−3.88	−2.04	−1.23
(NA)(HCl)(A)$(W)_2$ + W ⇌ (NA)(HCl)(A)$(W)_3$	−3.42	−1.85	−1.16

Table 2 DLPNO-CCSD(T)/CBS//DLPNO-CCSD(T)/CBS//ωB97X-D/6-31++G** Gibbs free energy changes (kcal mol^{-1}) associated with the formation and sequential hydration of two acid and dimethylamine trimers at atmospherically relevant temperatures and 1 atm pressure. Values from a[72] and b[86]

Cluster	216.65 K	273.15 K	298.15 K
SA + FA + DMA ⇌ (SA)(FA)(DMA)[a]	−26.08	−22.19	−20.48
(SA)(FA)(DMA) + W ⇌ (SA)(FA)(DMA)(W)[a]	−3.31	−1.57	−0.80
(SA)(FA)(DMA)(W) + W ⇌ (SA)(FA)(DMA)$(W)_2$[a]	−1.93	0.00	0.85
(SA)(FA)(DMA)$(W)_2$ + W ⇌ (SA)(FA)(DMA)$(W)_3$[a]	−1.21	0.50	1.25
SA + NA + DMA ⇌ (SA)(NA)(DMA)[a]	−25.19	−21.37	−19.69
(SA)(NA)(DMA) + W ⇌ (SA)(NA)(DMA)(W)[a]	−3.60	−1.66	−0.89
(SA)(NA)(DMA)(W) + W ⇌ (SA)(NA)(DMA)$(W)_2$[a]	−2.95	−1.24	−0.42
(SA)(NA)(DMA)$(W)_2$ + W ⇌ (SA)(NA)(DMA)$(W)_3$[a]	−2.35	−0.62	0.16
SA + HCl + DMA ⇌ (SA)(HCl)(DMA)[b]	−23.69	−20.10	−18.50
(SA)(HCl)(DMA) + W ⇌ (SA)(HCl)(DMA)(W)[b]	−3.96	−1.97	−1.09
(SA)(HCl)(DMA)(W) + W ⇌ (SA)(HCl)(DMA)$(W)_2$[b]	−2.75	−0.93	−0.13
(SA)(HCl)(DMA)$(W)_2$ + W ⇌ (SA)(HCl)(DMA)$(W)_3$[b]	−1.96	−0.24	0.52
FA + NA + DMA ⇌ (FA)(NA)(DMA)[a]	−15.58	−11.46	−9.65
(FA)(NA)(DMA) + W ⇌ (FA)(NA)(DMA)(W)[a]	−2.34	−0.71	0.00
(FA)(NA)(DMA)(W) + W ⇌ (FA)(NA)(DMA)$(W)_2$[a]	−1.42	0.34	1.12
(FA)(NA)(DMA)$(W)_2$ + W ⇌ (FA)(NA)(DMA)$(W)_3$[a]	−1.66	0.31	1.18
FA + HCl + DMA ⇌ (FA)(HCl)(DMA)[b]	−15.73	−12.09	−10.48
(FA)(HCl)(DMA) + W ⇌ (FA)(HCl)(DMA)(W)[b]	−2.68	−1.01	−0.28
(FA)(HCl)(DMA)(W) + W ⇌ (FA)(HCl)(DMA)$(W)_2$[b]	−2.23	−0.16	0.75
(FA)(HCl)(DMA)$(W)_2$ + W ⇌ (FA)(HCl)(DMA)$(W)_3$[b]	−2.78	−1.08	−0.33
NA + HCl + DMA ⇌ (NA)(HCl)(DMA)	−14.50	−10.92	−9.34
(NA)(HCl)(DMA) + W ⇌ (NA)(HCl)(DMA)(W)	−4.06	−2.42	−1.70
(NA)(HCl)(DMA)(W) + W ⇌ (NA)(HCl)(DMA)$(W)_2$	−2.96	−1.23	−0.47
(NA)(HCl)(DMA)$(W)_2$ + W ⇌ (NA)(HCl)(DMA)$(W)_3$	−2.86	−0.78	0.14

Waals forces where either the bond angle encompassing the hydrogen is less than 140° or the hydrogen-bond distance is greater than 2.2 Å. The top half of Fig. 2 illustrates the Gibbs free energy minima (298 K) for HCl bound to one of the other acids (SA, FA, or NA), the ammonia base, and 0–3 waters, while the bottom half of the figure shows how these combinations of

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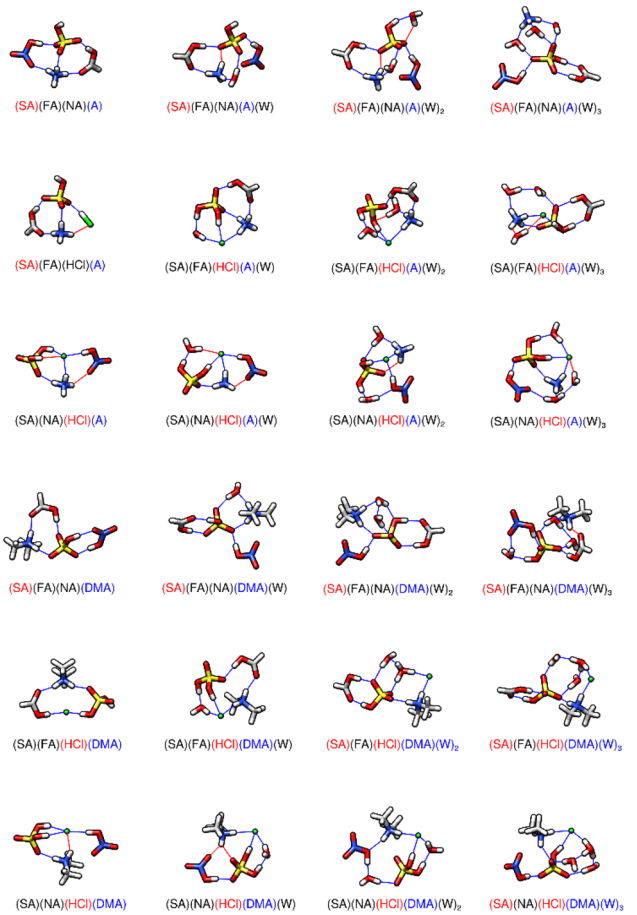

Fig. 3 DLPNO-CCSD(T)/CBS//ωB97X-D/6-31++G** minimum energy structures for the sequential hydration of three acids and one base. The molecule labels are colored according to charge as follows: blue = +1, black = 0, red = −1. Atoms are drawn in the following colors: hydrogen – white, carbon – grey, nitrogen – blue, oxygen – red, sulfur – yellow, chlorine – green.

acids are complexed with the DMA base and water. Unsurprisingly, as displayed in Table 1, two acid–one base trimers with SA have the most negative energies of formation. The two acid–ammonia trimer with the most positive energy of formation is (NA)(HCl)(A) (Table 1). This is likely due to the dry cluster only forming two hydrogen bonds (Fig. 2). However, this cluster has a very large first sequential hydration energy of −5.30 kcal-mol^{-1} at 217 K (Table 1), probably caused by a proton transfer and the addition of two new hydrogen bonds and one van der Waals interaction. The (NA)(HCl)(A)(W) cluster has the largest magnitude for the first hydration energy of any other complex that contains HCl (*vide infra*). The second sequential hydration energy is −3.88 kcal mol^{-1} at 217 K, which is the most negative second hydration energy of the two acid–ammonia trimers (Table 1). Similar to the one water cluster, HCl and A are in their charged forms with NA remaining neutral. (NA)(HCl)(A)$(W)_2$ has five hydrogen bonds and one van der Waals interaction, gaining only one hydrogen bond upon the addition of the second water (Fig. 2). For the two acid–DMA trimers, (NA)(HCl)(base) again has the most positive energy of formation but also has the most negative sequential hydration energies. (NA)(HCl)(DMA) has only two hydrogen bonds and one van der Waals interaction. Upon first hydration, only one new hydrogen bond is formed, with the DMA donating a hydrogen bond to the water. Additional water molecules add to the previous water(s), forming a ring between the DMA, HCl, and waters and leaving the NA on the outside of the cluster in every scenario. These results agree with previous results where clusters with HCl have more negative energies of sequential hydration than other clusters.[86]

3.2 Tetramers of three acids and one base

The structures and energies of formation and hydration for the tetramers of three acids and one base are shown in Fig. 3, Tables 3 & 4. When the singular base is A, the (SA)(NA)(HCl)(A) cluster has the most positive formation energy. This is consistent with the trimers, in which the cluster containing NA, HCl, and A was also the least stable. The (SA)(NA)(HCl)(A) dry cluster contains four hydrogen bonds, two van der Waals interactions, and a proton transfer from the HCl to the A (Fig. 3). Despite the dry cluster being the least stable (−11.05 kcal mol^{-1} at 298 K), the (SA)(NA)(HCl)(A)$(W)_{1-3}$ clusters have the most negative hydration energies. Due to these large negative hydration energies,

Table 3 DLPNO-CCSD(T)/CBS//ωB97X-D/6-31++G** Gibbs free energy changes (kcal mol-^{1} associated with the formation and sequential hydration of three acid and ammonia tetramers at atmospherically relevant temperatures and 1 atm pressure. Values from a[72] and b[86]

Cluster	216.65 K	273.15 K	298.15 K
SA + FA + NA + DMA ⇌ (SA)(FA)(NA)(DMA)[a]	−33.41	−27.50	−24.90
(SA)(FA)(NA)(DMA) + W ⇌ (SA)(FA)(NA)(DMA)(W)[a]	−2.39	−0.48	0.36
(SA)(FA)(NA)(DMA)(W) + W ⇌ (SA)(FA)(NA)(DMA)$(W)_2$[a]	−1.81	−0.01	0.78
(SA)(FA)(NA)(DMA)$(W)_2$ + W ⇌ (SA)(FA)(NA)(DMA)$(W)_3$[a]	−0.87	0.82	1.57
SA + FA + HCl + DMA ⇌ (SA)(FA)(HCl)(DMA)[b]	−28.92	−23.10	−20.59
(SA)(FA)(HCl)(DMA) + W ⇌ (SA)(FA)(HCl)(DMA)(W)[b]	−4.93	−3.06	−2.17
(SA)(FA)(HCl)(DMA)(W) + W ⇌ (SA)(FA)(HCl)(DMA)$(W)_2$[b]	−2.69	−0.85	−0.03
(SA)(FA)(HCl)(DMA)$(W)_2$ + W ⇌ (SA)(FA)(HCl)(DMA)$(W)_3$[b]	−2.04	−0.34	0.42
SA + NA + HCl + DMA ⇌ (SA)(NA)(HCl)(DMA)	−29.60	−24.16	−21.76
(SA)(NA)(HCl)(DMA) + W ⇌ (SA)(NA)(HCl)(DMA)(W)	−4.06	−1.99	−1.11
(SA)(NA)(HCl)(DMA)(W) + W ⇌ (SA)(NA)(HCl)(DMA)$(W)_2$	−3.29	−1.47	−0.62
(SA)(NA)(HCl)(DMA)$(W)_2$ + W ⇌ (SA)(NA)(HCl)(DMA)$(W)_3$	−2.43	−0.27	0.69

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Table 4 DLPNO-CCSD(T)/CBS//ωB97X-D/6-31++G** Gibbs free energy changes (kcal mol^{-1}) associated with the formation and sequential hydration of three acid and dimethylamine tetramers at atmospherically relevant temperatures and 1 atm pressure. Values from a[72] and b[86]

Cluster	216.65 K	273.15 K	298.15 K
SA + FA + NA + DMA ⇌ (SA)(FA)(NA)(DMA)[a]	−33.41	−27.50	−24.90
(SA)(FA)(NA)(DMA) + W ⇌ (SA)(FA)(NA)(DMA)(W)[a]	−2.39	−0.48	0.36
(SA)(FA)(NA)(DMA)(W) + W ⇌ (SA)(FA)(NA)(DMA)$(W)_2$[a]	−1.81	−0.01	0.78
(SA)(FA)(NA)(DMA)$(W)_2$ + W ⇌ (SA)(FA)(NA)(DMA)$(W)_3$[a]	−0.87	0.82	1.57
SA + FA + HCl + DMA ⇌ (SA)(FA)(HCl)(DMA)[b]	−28.92	−23.10	−20.59
(SA)(FA)(HCl)(DMA) + W ⇌ (SA)(FA)(HCl)(DMA)(W)[b]	−4.93	−3.06	−2.17
(SA)(FA)(HCl)(DMA)(W) + W ⇌ (SA)(FA)(HCl)(DMA)$(W)_2$[b]	−2.69	−0.85	−0.03
(SA)(FA)(HCl)(DMA)$(W)_2$ + W ⇌ (SA)(FA)(HCl)(DMA)$(W)_3$[b]	−2.04	−0.34	0.42
SA + NA + HCl + DMA ⇌ (SA)(NA)(HCl)(DMA)	−29.60	−24.16	−21.76
(SA)(NA)(HCl)(DMA) + W ⇌ (SA)(NA)(HCl)(DMA)(W)	−4.06	−1.99	−1.11
(SA)(NA)(HCl)(DMA)(W) + W ⇌ (SA)(NA)(HCl)(DMA)$(W)_2$	−3.29	−1.47	−0.62
(SA)(NA)(HCl)(DMA)$(W)_2$ + W ⇌ (SA)(NA)(HCl)(DMA)$(W)_3$	−2.43	−0.27	0.69

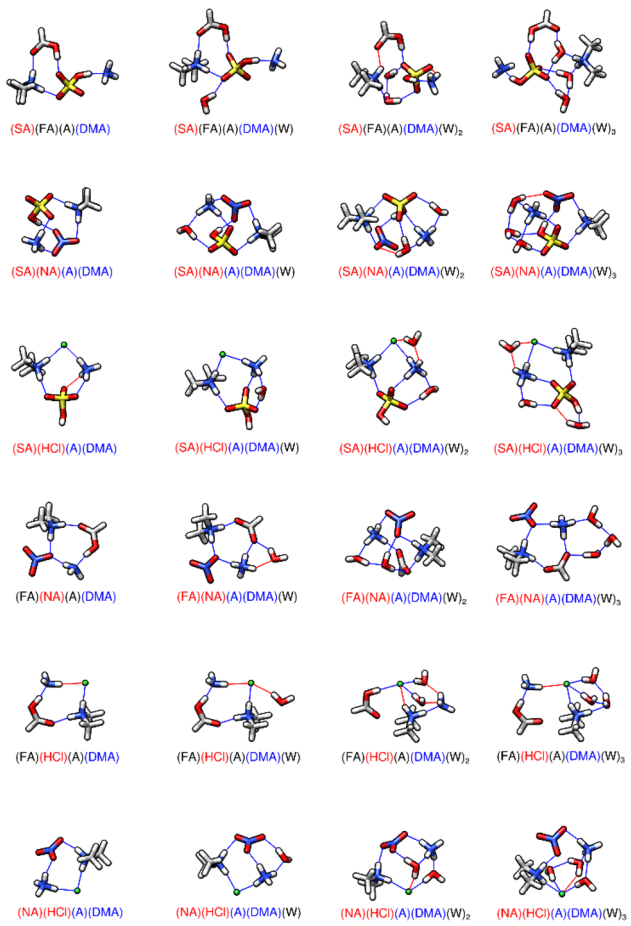

Fig. 4 DLPNO-CCSD(T)/CBS//ωB97X-D/6-31++G** minimum energy structures for the sequential hydration of two acids and two bases. The molecule labels are colored according to charge as follows: blue = +1, black = 0, red = −1. Atoms are drawn in the following colors: hydrogen – white, carbon – grey, nitrogen – blue, oxygen – red, sulfur – yellow, chlorine – green.

(SA)(FA)(NA)(A)$(W)_3$ is only 1.27 kcal mol^{-1} more stable than (SA)(NA)(HCl)(A)$(W)_3$ at 298 K, as compared to their dry clusters, which have a free energy difference of 5.23 kcal mol^{-1} (Table 3). This emphasizes the hydrating power of HCl. When the base is DMA, all dry clusters have formation energies averaging roughly 9.5 kcal mol^{-1} more negative than their A counterparts (Table 4: the exact differences are −8.64 for SA–FA–NA; −8.72 for SA–FA–HCl; −10.84 for SA–NA–HCl at 217 K). While both (SA)(FA)(NA)(base) clusters remain the most stable in the dry state, the (SA)(FA)(HCl)(DMA) is the least stable. All dry clusters are di-ionic and contain five hydrogen bonds, except for (SA)(NA)(HCl)(DMA), which contains four hydrogen bonds and one van der Waals interaction (Fig. 3). The clusters containing HCl have more negative hydration energies than (SA)(FA)(NA)(DMA)$(W)_{1–3}$, as seen in Table 4. In fact, these hydration energies are so much lower for the clusters containing HCl that (SA)(FA)(NA)(DMA)$(W)_3$ is the least stable of the tertiary hydrated three acid–DMA clusters, despite being the most stable in the dry state. This is likely in part caused by the di-ionic nature of the (SA)(FA)(NA)(DMA)$(W)_3$ cluster, which only has eight hydrogen bonds and one van der Waals interaction, whereas the other tertiary hydrated three acid–DMA clusters are tetra-ionic. These structures have nine hydrogen bonds and reveal a proton transfer to water forming a hydronium ion (Fig. 3). The more negative sequential hydration values for (SA)(NA)(HCl)(base)$(W)_{2–3}$ hint that three stronger acids and one base can hydrate more effectively than two strong acids, one weak acid, and one base. This is illustrated by comparing the Gibbs free energy values for hydration of the second and third waters in Tables 3 and 4. The values for hydration of (SA)(NA)(HCl)(A)$(W)_2$ and (SA)(NA)(HCl)(A)$(W)_3$ are −3.80 and −3.53 kcal mol^{-1} at 217 K, respectively (Table 3). Substituting FA for NA in these clusters result in the next lowest hydration energies (−2.88 for W_2; −3.05 for W_3 at 217 K), and the most positive hydration energies are for the (SA)(FA)(NA)(A)$(W)_{2–3}$ clusters. Table 4 shows that substituting DMA for W has the same exact trends, with (SA)(NA)(HCl)(DMA)$(W)_{2–3}$ having the most negative hydration free energies (−3.29 and −2.43 kcal mol^{-1} at 217 K), followed by the clusters where FA substitutes for NA. While this may be related to structural factors, it may also highlight the effects that three acids together have on a singular base. It is possible that acid strength is more important as the clusters are hydrated, or it may reaffirm the stabilizing effects that HCl has in hydrating clusters.[86]

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Table 5 DLPNO-CCSD(T)/CBS//ωB97X-D/6-31++G** Gibbs free energy changes (kcal mol^{-1}) associated with the formation and sequential hydration of two acid and two base tetramers at atmospherically relevant temperatures and 1 atm pressure. Values from a[72] and b[86]

Cluster	216.65 K	273.15 K	298.15 K
SA + FA + A + DMA ⇌ (SA)(FA)(A)(DMA)[a]	−30.63	−24.98	−22.49
(SA)(FA)(A)(DMA) + W ⇌ (SA)(FA)(A)(DMA)(W)[a]	−2.25	−0.50	0.27
(SA)(FA)(A)(DMA)(W) + W ⇌ (SA)(FA)(A)(DMA)(W)$_2$[a]	−1.07	1.10	1.99
(SA)(FA)(A)(DMA)(W)$_2$ + W ⇌ (SA)(FA)(A)(DMA)(W)$_3$[a]	−2.05	−0.48	0.28
SA + NA + A + DMA ⇌ (SA)(NA)(A)(DMA)[a]	−33.31	−27.23	−24.55
(SA)(NA)(A)(DMA) + W ⇌ (SA)(NA)(A)(DMA)(W)[a]	−6.79	−4.84	−3.99
(SA)(NA)(A)(DMA)(W) + W ⇌ (SA)(NA)(A)(DMA)(W)$_2$[a]	−1.60	0.48	1.40
(SA)(NA)(A)(DMA)(W)$_2$ + W ⇌ (SA)(NA)(A)(DMA)(W)$_3$[a]	−1.62	0.22	1.04
SA + HCl + A + DMA ⇌ (SA)(HCl)(A)(DMA)[b]	−36.33	−30.87	−28.45
(SA)(HCl)(A)(DMA) + W ⇌ (SA)(HCl)(A)(DMA)(W)[b]	−4.26	−2.45	−1.66
(SA)(HCl)(A)(DMA)(W) + W ⇌ (SA)(HCl)(A)(DMA)(W)$_2$[b]	−3.18	−1.56	−0.84
(SA)(HCl)(A)(DMA)(W)$_2$ + W ⇌ (SA)(HCl)(A)(DMA)(W)$_3$[b]	−2.14	−0.36	0.42
FA + NA + A + DMA ⇌ (FA)(NA)(A)(DMA)[a]	−19.35	−13.94	−11.57
(FA)(NA)(A)(DMA) + W ⇌ (FA)(NA)(A)(DMA)(W)[a]	−2.16	−0.32	0.50
(FA)(NA)(A)(DMA)(W) + W ⇌ (FA)(NA)(A)(DMA)(W)$_2$[a]	−1.82	0.38	1.36
(FA)(NA)(A)(DMA)(W)$_2$ + W ⇌ (FA)(NA)(A)(DMA)(W)$_3$[a]	−2.73	−1.23	−0.66
FA + HCl + A + DMA ⇌ (FA)(HCl)(A)(DMA)[b]	−17.85	−12.63	−10.32
(FA)(HCl)(A)(DMA) + W ⇌ (FA)(HCl)(A)(DMA)(W)[b]	−2.50	−0.79	−0.04
(FA)(HCl)(A)(DMA)(W) + W ⇌ (FA)(HCl)(A)(DMA)(W)$_2$[b]	−2.16	−0.10	0.81
(FA)(HCl)(A)(DMA)(W)$_2$ + W ⇌ (FA)(HCl)(A)(DMA)(W)$_3$[b]	−2.15	−0.74	−0.12
NA + HCl + A + DMA ⇌ (NA)(HCl)(A)(DMA)	−24.37	−18.73	−16.25
(NA)(HCl)(A)(DMA) + W ⇌ (NA)(HCl)(A)(DMA)(W)	−3.73	−1.97	−1.20
(NA)(HCl)(A)(DMA)(W) + W ⇌ (NA)(HCl)(A)(DMA)(W)$_2$	−2.49	−0.62	0.21
(NA)(HCl)(A)(DMA)(W)$_2$ + W ⇌ (NA)(HCl)(A)(DMA)(W)$_3$	0.08	1.75	2.49

3.3 Tetramers of two acids and two bases

The structures and energies of formation and hydration for the tetramers of two acids and two bases are displayed in Fig. 4 and Table 5. The left-hand side of the figure reveals that the dry clusters of the stronger acids are all tetra-ionic, but if FA replaces one of the stronger acids, the resulting complex is di-ionic. (SA)(HCl)(A)(DMA) is the most negative two acid–two base tetramer with a $\Delta G°$ value of −36.33 kcal mol^{-1} at 217 K (Table 5). When SA is replaced with NA to form (NA)(HCl)(A)(DMA), the $\Delta G°$ of formation becomes more positive by 11.96 kcal mol^{-1}. Both (NA)(HCl)(A)(DMA) and (SA)(HCl)(A)(DMA) have four hydrogen bonds, with (SA)(HCl)(A)(DMA) also having one van der Waals interaction. The large energy difference between the two tetramers displays the driving effects of SA. However, if FA replaces NA, the $\Delta G°$ of binding is more positive than for (NA)(HCl)(A)(DMA) by 6.52 kcal mol^{-1}. In fact, (FA)(HCl)(A)(DMA) is the least stable two acid–two base tetramer, with a $\Delta G°$ of formation of −17.85 kcal mol^{-1} at 217 K (Table 5). (FA)(HCl)(A)(DMA) has four hydrogen bonds, like the others, but (FA)(HCl)(A)(DMA) is di-ionic with a proton exchange between HCl and DMA (Fig. 4). The two acid–two base tetramers illustrate the competitive factors that contribute to the subtle stability of a pre-nucleation cluster: hydrogen bonding topology, protonation states, acid strength, and the driving effects of specific monomers. Upon hydration, the trimers with HCl have on average more negative energies of sequential hydration than those without. The primary exception to this pattern is (SA)(NA)(A)(DMA)(W), which has been previously noted for its special stability upon hydration.[72] (SA)(HCl)(A)(DMA)(W) has the second most negative energy of first hydration, −4.26 kcal mol^{-1} at 217 K, followed closely by (NA)(HCl)(A)(DMA)(W) at −3.73 kcal mol^{-1} (Table 5). Both (SA)(HCl)(A)(DMA)(W) and (NA)(HCl)(A)(DMA)(W) have six hydrogen bonds and are tetra-ionic (Fig. 4). The 0.53 kcal mol^{-1} difference in first sequential hydration energy is likely a result of subtle hydrogen bonding topology differences. (SA)(HCl)(A)(DMA)(W)$_2$ has the most negative second sequential hydration energy, followed again by (NA)(HCl)(A)(DMA)(W)$_2$. Upon the third sequential hydration, all but one of the two acid–two base tetramers seem to have very similar energies, with no clear preference between clusters with or without HCl. Somewhat surprisingly, (NA)(HCl)(A)(DMA)(W)$_3$ has the most positive third sequential hydration energy, with positive $\Delta G°$ values at atmospherically relevant temperatures. However, examining the structures reveals this is because the (NA)(HCl)(A)(DMA)(W)$_3$ complex has only one more hydrogen bond than (NA)(HCl)(A)(DMA)(W)$_2$ (Fig. 4). Nitric acid is a much worse contributor to hydrogen bonds in this system compared to sulfuric acid. The two acid–two base cluster results reaffirm the idea that hydrogen bonding topology is influential for cluster stability and energetics.

3.4 Pentamer of three acids and two bases

The structures and energies of formation and hydration for the pentamers of three acids and two bases are shown in Fig. 5 and Table 6. At 217 K, the $\Delta G°$ of formation for the dry pentamer (SA)(NA)(HCl)(A)(DMA) is 0.34 kcal mol^{-1} more positive than (SA)(FA)(HCl)(A)(DMA) (Table 6). Both dry pentamers are tetra-ionic, with SA and HCl donating protons to A and DMA. Additionally, both dry structures have six hydrogen bonds and

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display acid–base bridging between the charged species. A structural feature of (SA)(FA)(HCl)(A)(DMA) that may result in its slightly more negative energy is the presence of the SA–FA dimer, which Harold *et al.* showed to have exceptional stability (Fig. 5).65 (SA)(FA)(NA)(A)(DMA) is also tetra-ionic with six hydrogen bonds and has a ΔG° of binding 3.18 kcal mol^{-1} more positive than (SA)(NA)(HCl)(A)(DMA) at 217 K (Table 6). The dry three acid–two base pentamers with HCl have more negative ΔG° binding values than the pentamer without HCl. When the pentamers are hydrated with one water, the pentamers with HCl gain two hydrogen bonds, while (SA)(FA)(NA)(A)(DMA)(W) only gains one hydrogen bond and one van der Waals interaction. However (SA)(NA)(HCl)(A)(DMA) has the most positive energy of first sequential hydration by 1.52 kcal mol^{-1} relative to when FA replaces HCl (Table 6). (SA)(FA)(HCl)(A)(DMA)(W) has the most negative first sequential hydration energy, −4.35 kcal mol^{-1}, nearly double that of (SA)(FA)(HCl)(A)(DMA)(W). In (SA)(NA)(HCl)(A)(DMA)(W), NA is loosely bound to the rest of the cluster through one hydrogen bond accepted by SA. The other two primary hydrated pentamers form the SA–FA dimer, which seems to have a stabilizing effect on the first hydration energy of the three acid–two base pentamers, in addition to hydrogen bonding topology. (SA)(FA)(HCl)(A)(DMA)(W)2 has the most positive second sequential hydration energy, despite forming one more hydrogen bond than (SA)(NA)(HCl)(A)(DMA)(W)2 and incorporating the SA–FA dimer. The (SA)(FA)(HCl)(A)(DMA)(W)3 cluster also maintains the SA–FA dimer, as well as gaining one hydrogen bond and one van der Waals interaction, yet has no additional stabilization at 217 K. The only tertiary sequential hydration that has negative ΔG° values is for (SA)(NA)(HCl)(A)(DMA)(W)3, which has a total of eleven hydrogen bonds and is the most stable three water pentamer. The three acid–two base pentamers demonstrate the stabilizing power of nitric acid combined with hydrochloric acid when sufficient bases are present.

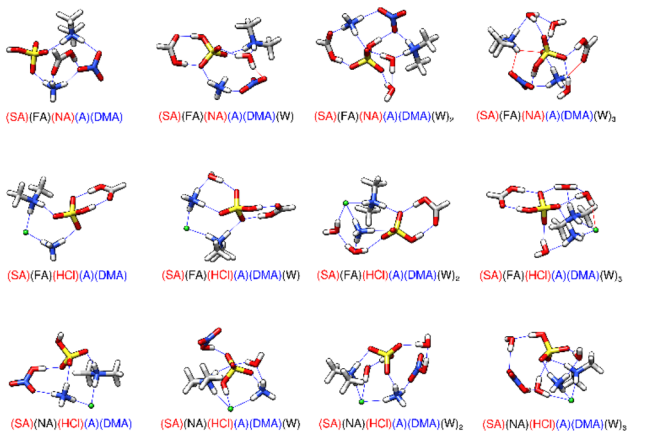

Fig. 5 DLPNO-CCSD(T)/CBS//ωB97X-D/6-31++G** minimum energy structures for the sequential hydration of three acids and two bases. The molecule labels are colored according to charge as follows: blue = +1, black = 0, red = −1. Atoms are drawn in the following colors: hydrogen – white, carbon – grey, nitrogen – blue, oxygen – red, sulfur – yellow, chlorine – green.

3.5 Equilibrium concentrations and pathways of formation

Equilibrium constants were calculated using the ΔG° values of the cluster at 216.65 and 298.15 K, corresponding to the top and bottom of the troposphere. Systems of equations adapted from Odbadrakh *et al.* were then used to calculate the equilibrium concentrations,[57,72] assuming a closed system consisting of SA, NA, FA, HCl, A, DMA, and three W. In terms of kinetics, the ratio between formation and evaporation reactions is accounted for by the equilibrium constants. Other kinetic factors, such as the timescale to reach equilibrium and potential scavenging or sink reactions will only be important at very low vapor concentrations. We chose a water concentration of 7.7×10^{17} cm^{-3} at 298 K and 9.9×10^{14} cm^{-3} at 217 K, which corresponds to 100% humidity at the bottom and top of the troposphere.[5] Initial starting concentrations of the monomers were 5×10^{7} cm^{-3} for SA, 9.8×10^{10} cm^{-3} for NA, 1×10^{9} cm^{-3} for HCl, 2×10^{11} cm^{-3} for A, and 2×10^{9} cm^{-3} for DMA at 298 K, which are relevant for each of the monomers over inland and urban areas.[5,30,118–125] At 217 K, the monomer concentrations are decreased by three orders of magnitude to account for atmospheric thinning at the top of the troposphere. This is a rough estimate based on the three orders of decrease in concentration of water, and we resort to this approximation since experimental concentrations of

Table 6 DLPNO-CCSD(T)/CBS//ωB97X-D/6-31++G** Gibbs free energy changes (kcal mol^{-1}) associated with the formation and sequential hydration of three acid and two base pentamers at atmospherically relevant temperatures and 1 atm pressure. Values from a[72] and b[86]

Cluster	216.65 K	273.15 K	298.15 K
SA + FA + NA + A + DMA ⇌ (SA)(FA)(NA)(A)(DMA)[a]	−39.56	−31.53	−28.00
(SA)(FA)(NA)(A)(DMA) + W ⇌ (SA)(FA)(NA)(A)(DMA)(W)[a]	−4.13	3.09	4.03
(SA)(FA)(NA)(A)(DMA)(W) + W ⇌ (SA)(FA)(NA)(A)(DMA)(W)$_2$[a]	−4.07	−1.54	−0.64
(SA)(FA)(NA)(A)(DMA)(W)$_2$ + W ⇌ (SA)(FA)(NA)(A)(DMA)(W)$_3$[a]	1.16	3.69	4.58
SA + FA + HCl + A + DMA ⇌ (SA)(FA)(HCl)(A)(DMA)[b]	−43.08	−35.34	−32.02
(SA)(FA)(HCl)(A)(DMA) + W ⇌ (SA)(FA)(HCl)(A)(DMA)(W)[b]	−4.35	−2.67	−1.82
(SA)(FA)(HCl)(A)(DMA)(W) + W ⇌ (SA)(FA)(HCl)(A)(DMA)(W)$_2$[b]	−2.85	−0.85	0.04
(SA)(FA)(HCl)(A)(DMA)(W)$_2$ + W ⇌ (SA)(FA)(HCl)(A)(DMA)(W)$_3$[b]	0.00	2.07	2.86
SA + NA + HCl + A + DMA ⇌ (SA)(NA)(HCl)(A)(DMA)	−42.74	−35.29	−32.00
(SA)(NA)(HCl)(A)(DMA) + W ⇌ (SA)(NA)(HCl)(A)(DMA)(W)	−2.61	−0.35	0.66
(SA)(NA)(HCl)(A)(DMA)(W) + W ⇌ (SA)(NA)(HCl)(A)(DMA)(W)$_2$	−3.67	−1.73	−0.86
(SA)(NA)(HCl)(A)(DMA)(W)$_2$ + W ⇌ (SA)(NA)(HCl)(A)(DMA)(W)$_3$	−2.60	−0.80	−0.01

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these monomers in the upper troposphere are difficult to measure. The calculated equilibrium concentrations of atmospherically relevant clusters are shown in Table 7, where we have defined atmospherically relevant as equilibrium concentrations above one cm^{-3} at either temperature. Among the atmospherically relevant dimers, the (SA)(DMA) dimer is the most abundant. The NA–base dimers are present in higher concentrations, for both dry and hydrated conditions, than the HCl–base dimers and the FA–base dimers discussed previously.[72,86] The largest dry cluster in this simulation that is atmospherically relevant is (SA)(HCl)(A)(DMA), which is the most stable dry two acid–two base tetramer by 3 kcal mol^{-1} at 217 K. In comparison, the largest dry cluster with NA that is atmospherically relevant is the (SA)(NA)(DMA) trimer, indicating that NA might evaporate out of the clusters before they grow too large. When the system is hydrated, clusters with NA have larger equilibrium constants than those with HCl, similar to the dimer results. The largest hydrated cluster with HCl is $(HCl)(DMA)(W)_2$, and the largest hydrated cluster with NA is (SA)(NA)(A)(DMA)(W). As previously noted (SA)(NA)(A)(DMA)(W) is exceptionally stable, explaining its appearance as the only hydrated tetramer.[72] While clusters with HCl have more favorable Gibbs free energies of sequential hydration than those with NA, clusters with NA are more atmospherically relevant in their hydrated forms, likely due to their higher initial concentrations. This is because in our simulation the initial concentration of NA is about two orders of magnitude higher than the concentration of HCl. If the concentrations were the same, the hydrated HCl clusters would be more important. Under our simulation conditions there are no clusters containing both NA and HCl that have concentrations above one cm^{-3} at 217 K or 298 K. However, investigating the growth pathways to form the dry (SA)(NA)(HCl)(A)(DMA) pentamer can reveal driving effects of the different monomers. Therefore, in Table 8, we show the optimal growth pathways for the dry (SA)(NA)(HCl)(A)(DMA) pentamer using the most atmospherically relevant paths at 217 K and 298 K. The optimal pathway for growth begins with (SA)(DMA) at both temperatures. SA and DMA are both known to be strong drivers of nucleation, which will promote cluster growth. NA is added next, which has stabilizing effects.[72] Ammonia and finally HCl are added to reach the dry pentamer. This mechanism is consistent with the usual alternate addition of acid and bases for other acid–base systems and leads to an approximately 1 : 1 ratio of acid to bases in the growing clusters. The second-best pathway starts with (NA)(DMA) and then adds SA. Therefore, the first- and second-best pathways pass through the same trimer, (SA)(NA)(DMA), which is atmospherically relevant at both temperatures. The only difference between the first- and second-best pathways is the acid in the starting dimer. This discrepancy is likely due to the balance between thermodynamics and mass balance, as (SA)(DMA) is thermodynamically favored but NA is present in a much higher concentration than SA. In both of these pathways, HCl is the last molecule to be added to the cluster, likely because HCl has a much greater stabilizing effect on hydrated clusters than dry clusters. The third-best pathway also starts with (SA)(DMA), followed by the addition of HCl. (SA)(HCl)(DMA) is barely atmospherically relevant, present only at 298 K and in a very low concentration (Table 7). Ammonia is added to form the tetramer (SA)(HCl)(A)(DMA) and NA is added last. When the three systems (SA)(FA)(NA)(A)(DMA),[72,86] (SA)(FA)(HCl)(A)(DMA),[72,86] and (SA)(NA)(HCl)(A)(DMA) are compared, SA and DMA stand out as strong drivers of prenucleation due to their higher equilibrium concentrations. In this particular system, NA appears more in

Table 7 Equilibrium concentrations of clusters that form at more than 1 cm^{-3} at 217 or 298 K. Initial concentrations of the monomers at 217 K were: SA = 5.00×10^4, NA = 9.80×10^7, HCl = 1.00×10^6, A = 2.00×10^8, DMA = 2.00×10^6 cm^{-3}. Initial concentrations of the monomers at 298 K were: SA = 5.00×10^7, NA = 9.80×10^{10}, HCl = 1.00×10^9, A = 2.00×10^{11}, DMA = 2.00×10^9 cm^{-3}. The water concentration is 9.90×10^{17} cm^{-3} at 217 K and 7.70×10^{14} cm^{-3} at 298 K. See text for details

Cluster	216.65 K	298.15 K
SA	2.96×10^1	1.48×10^7
NA	9.77×10^1	9.63×10^{10}
HCl	1.00×10^6	9.99×10^8
A	2.00×10^8	2.00×10^{11}
DMA	1.95×10^6	1.98×10^9
(SA)(NA)	6.50×10^{-6}	2.10
(SA)(A)	9.50×10^{-2}	5.99×10^3
(SA)(DMA)	2.53×10^4	1.14×10^7
(NA)(A)	2.69×10^2	2.16×10^5
(NA)(DMA)	1.89×10^2	2.23×10^4
(HCl)(A)	2.26×10^{-4}	3.73
(HCl)(DMA)	4.39×10^{-3}	4.24
(SA)(NA)(DMA)	1.26×10^2	1.26×10^3
(SA)(HCl)(DMA)	3.97×10^{-2}	1.75
(SA)(A)(DMA)	2.51×10^{-3}	1.05
(SA)(HCl)(A)(DMA)	1.32	2.80×10^{-1}
(SA)(W)	2.48×10^1	1.50×10^7
$(SA)(W)_2$	3.41	2.90×10^6
$(SA)(W)_3$	7.48×10^{-2}	1.45×10^5
(NA)(W)	3.19×10^5	1.73×10^9
$(NA)(W)_2$	3.04×10^2	1.33×10^7
$(NA)(W)_3$	3.11×10^{-2}	1.50×10^4
(HCl)(W)	1.24×10^1	7.01×10^5
$(HCl)(W)_2$	3.79×10^{-4}	3.39×10^2
$(HCl)(W)_3$	7.60×10^{-7}	3.61
(A)(W)	6.57×10^3	3.79×10^8
$(A)(W)_2$	5.61×10^{-2}	4.52×10^4
$(A)(W)_3$	1.88×10^{-5}	1.03×10^2
(DMA)(W)	1.86×10^2	5.01×10^6
$(DMA)(W)_2$	1.04×10^{-2}	3.76×10^3
$(DMA)(W)_3$	2.47×10^{-6}	5.54
(SA)(NA)(W)	7.06×10^{-6}	2.13
(SA)(A)(W)	6.65×10^{-3}	5.61×10^2
$(SA)(A)(W)_2$	9.05×10^{-1}	1.29×10^4
$(SA)(A)(W)_3$	3.82×10^{-4}	6.31×10^1
(SA)(DMA)(W)	2.38×10^4	5.42×10^6
$(SA)(DMA)(W)_2$	7.38×10^2	3.06×10^5
$(SA)(DMA)(W)_3$	8.04×10^1	4.30×10^4
(NA)(A)(W)	9.48×10^{-3}	1.28×10^2
(NA)(DMA)(W)	6.73	1.42×10^3
$(NA)(DMA)(W)_2$	1.38×10^{-1}	6.54×10^1
$(NA)(DMA)(W)_3$	1.25×10^{-3}	3.40
(HCl)(DMA)(W)	2.48×10^{-2}	2.06×10^1
$(HCl)(DMA)(W)_2$	5.95×10^{-3}	5.60
(SA)(NA)(DMA)(W)	1.58×10^1	1.77×10^2
$(SA)(NA)(DMA)(W)_2$	4.37×10^{-1}	1.12×10^1
(SA)(NA)(A)(DMA)(W)	2.39×10^1	9.81×10^{-1}

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Table 8 Optimal pathways for growing the dry pentamer at 1 atm pressure according to equilibrium concentrations calculated using initial concentrations of $[SA]_0 = 5.00 \times 10^7$, $[NA]_0 = 9.80 \times 10^{10}$, $[HCl]_0 = 1.00 \times 10^9$, $[A]_0 = 2.00 \times 10^{11}$, and $[DMA]_0 = 2.00 \times 10^9$ cm^{-3} at 298 K. Concentrations were decreased by 3 orders of magnitude at 217 K. See text for details

	216.65 K	298.15 K
Optimal pathway	SA + DMA ⇌ (SA)(DMA) + NA ⇌ (SA)(NA)(DMA) + A ⇌ (SA)(NA)(A)(DMA) + HCl ⇌ (SA)(NA)(HCl)(A)(DMA)	SA + DMA ⇌ (SA)(DMA) + NA ⇌ (SA)(NA)(DMA) + A ⇌ (SA)(NA)(A)(DMA) + HCl ⇌ (SA)(NA)(HCl)(A)(DMA)
2nd best pathway	NA + DMA ⇌ (NA)(DMA) + SA ⇌ (SA)(NA)(DMA) + A ⇌ (SA)(NA)(A)(DMA) + HCl ⇌ (SA)(NA)(HCl)(A)(DMA)	NA + DMA ⇌ (NA)(DMA) + SA ⇌ (SA)(NA)(DMA) + A ⇌ (SA)(NA)(A)(DMA) + HCl ⇌ (SA)(NA)(HCl)(A)(DMA)
3rd best pathway	SA + DMA ⇌ (SA)(DMA) + HCl ⇌ (SA)(HCl)(DMA) + A ⇌ (SA)(HCl)(A)(DMA) + NA ⇌ (SA)(NA)(HCl)(A)(DMA)	SA + DMA ⇌ (SA)(DMA) + HCl ⇌ (SA)(HCl)(DMA) + A ⇌ (SA)(HCl)(A)(DMA) + NA ⇌ (SA)(NA)(HCl)(A)(DMA)

both dry and hydrated clusters than HCl does, likely caused by a higher initial concentration of NA than for HCl in our calculations. Many subtleties are at play in the beginning stages of pre-nucleation and these results cannot be attributed to only one factor. Hydrogen bonding topology and complex structural interactions appear to play an important role in structure energetics and atmospheric relevance.

4 Conclusions

For this system, sulfuric acid, nitric acid, hydrochloric acid, ammonia, dimethylamine, and three waters, we find that nitric acid is not stabilizing when paired with HCl in two acid–one base trimers. On the other hand, with two bases and an SA present, nitric acid is stabilizing when paired with HCl in the formation of three acid–two base pentamers. The charge on the monomers within clusters can be influential in the energetics, as the di-ionic and tetra-ionic two acid–two base tetramers illustrate. Acid strength is clearly important for hydrated clusters such as the three acid–one base tetramers. Complexes containing HCl tend to have more negative Gibbs free energies of hydration than those without HCl. Because deprotonated HCl simply leaves the chloride ion as a binding partner, clusters containing HCl tend to form more symmetric clusters, especially with water. This may explain the more negative Gibbs free energies of hydration for clusters containing HCl. This research adds to the body of work that illustrates that, depending on the system being studied, the acid/base strength of the monomers and the hydrogen bond topology of the resulting clusters have a subtle interplay that determines which cluster is most stable. In particular, sulfuric acid and dimethylamine excel at forming prenucleation complexes, and SA's ability to attract and bind additional molecules to the cluster drives growth. Other acids can't participate in as many hydrogen bonding events as bases add to a growing cluster.

Author contributions

The manuscript was written through contributions of all authors. All authors have given approval to the final version of the manuscript.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

Funding for this work was provided by grants CHE-16626238, CHE-1903871, and CHE-2018427 from the National Science Foundation (GCS), the Arnold and Mabel Beckman Foundation Beckman Scholar Award (CJB), and the Barry M. Goldwater Scholarship (CJB). High-performance computing resources were provided by the Research Corporation for Science Advancement (27446) and the MERCURY Consortium (https://www.mercuryconsortium.org/).[126,127] Molecular graphics and analyses performed with UCSF Chimera, developed by the Resource for Biocomputing, Visualization, and Informatics at the University of California, San Francisco, with support from NIH P41-GM103311. We thank the reviewers for helpful comments.

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