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Power-to-liquid *via* synthesis of methanol, DME or Fischer–Tropsch-fuels: a review

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The conversion of H_2 and CO_2 to liquid fuels *via* Power-to-Liquid (PtL) processes is gaining attention. With their higher energy densities compared to gases, the use of synthetic liquid fuels is particularly interesting in hard-to-abate sectors for which decarbonisation is difficult. However, PtL poses new challenges for the synthesis: away from syngas-based, continuously run, large-scale plants towards more flexible, small-scale concepts with direct CO_2-utilisation. This review provides an overview of state of the art synthesis technologies as well as current developments and pilot plants for the most prominent PtL routes for methanol, DME and Fischer–Tropsch-fuels. It should serve as a benchmark for future concepts, guide researchers in their process development and allow a technological evaluation of alternative reactor designs. In the case of power-to-methanol and power-to-FT-fuels, several pilot plants have been realised and the first commercial scale plants are planned or already in operation. In comparison power-to-DME is much less investigated and in an earlier stage of development. For methanol the direct CO_2 hydrogenation offers advantages through less by-product formation and lower heat development. However, increased water formation and lower equilibrium conversion necessitate new catalysts and reactor designs. While DME synthesis offers benefits with regards to energy efficiency, operational experience from laboratory tests and pilot plants is still missing. Furthermore, four major process routes for power-to-DME are possible, requiring additional research to determine the optimal concept. In the case of Fischer–Tropsch synthesis, catalysts for direct CO_2 utilisation are still in an early stage. Consequently, today's Fischer–Tropsch-based PtL requires a shift to syngas, benefiting from advances in co-electrolysis and reverse water-gas shift reactor design.

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Broader context

In their effort to achieve the goals set by the Paris Climate agreement, governments worldwide have facilitated the expansion of renewable energy sources and increased research funding for technologies such as solar and wind power, energy storage or carbon capture. At the same time, particularly the mobility sector is growing strongly, which results in increased greenhouse gas emissions and for many other sectors (*e.g.* chemical industry) a path towards emission reduction is not in sight. Power-to-Liquid (PtL) processes are seen as a promising solution to tackle these challenges. Using renewable electricity to produce liquid fuels with high energy densities, PtL can offer carbon neutral fuels for the mobility sector and new production routes for the chemical industry. Furthermore, PtL can contribute to grid stability and act as a long-term energy storage solution. In this review the most prominent PtL routes are considered. The review focuses on the synthesis step of each product and compares the state of the art with current research trends in the PtL context. In addition, an overview of PtL pilot plants is given and economic aspects are discussed. The goal is to identify research gaps and promote development steps towards commercialisation of PtL processes.

1. Introduction

Reducing greenhouse gas emissions has become a central target of worldwide energy policies. The Paris Climate agreement in 2015 set the aim of limiting global warming to below 2 °C compared to pre-industrial levels.[1] An increase of renewable energies, especially wind and solar power, represents an important component to realise this ambitious target. Their fluctuating feed-in results in an increased flexibility requirement of the energy system to balance supply and demand. Frontrunners like Denmark already face excess power production with wind generation supplying 130% of the load within a 24-hour window, giving an outlook for future renewable energy systems.[2] Large-scale short- and long-term energy storages will play a crucial role

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besides other flexibility options like demand side management to ensure electricity supply. Power-to-Gas (PtG) is often discussed as a potential long-term storage option by linking the power grid with the gas grid.[3] Several reviews concerning commercial methanation technologies,[4] technological and economic reviews of PtG-routes,[5] comparison of different renewable SNG pathways[6] and pilot plants[7–9] have been published. In addition entire books on the topic are available.[10–12] However, the conversion of electrolytic H_2 and CO_2 into liquid fuels (Power-to-Liquid, PtL) gains rising attention as a storage and flexibility option as well as carbon capture and utilisation (CCU) technology.

Fig. 1 shows that compared to Li-ion batteries and PtG products (hydrogen or methane), PtL products have the advantage of higher energy densities both in terms of weight and volume, making PtL processes especially interesting for hard-to-abate sectors *e.g.* aviation, shipping or heavy goods traffic. Moreover many PtL products represent a value added product, which otherwise has to be produced from natural gas (or coal). Consequently, PtL leads to a coupling of the electricity sector with the mobility sector and chemical industry. It enables excess renewable power from the electricity sector to be taken up and substitute fossil fuels in the mobility sector and chemical industry, resulting in a reduction of greenhouse gas emissions. The basic process chain of PtL as discussed within this review is shown in Fig. 2. Main process steps are: water electrolysis, carbon capture, synthesis and product upgrade.

Until now the commercial application of PtL processes is limited due to the remaining technological challenges regarding major process steps *e.g.* higher electrolysis efficiencies or improvements in carbon capture methods. In addition experience

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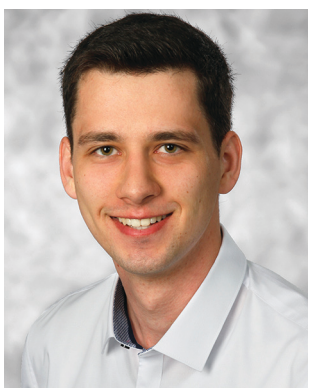

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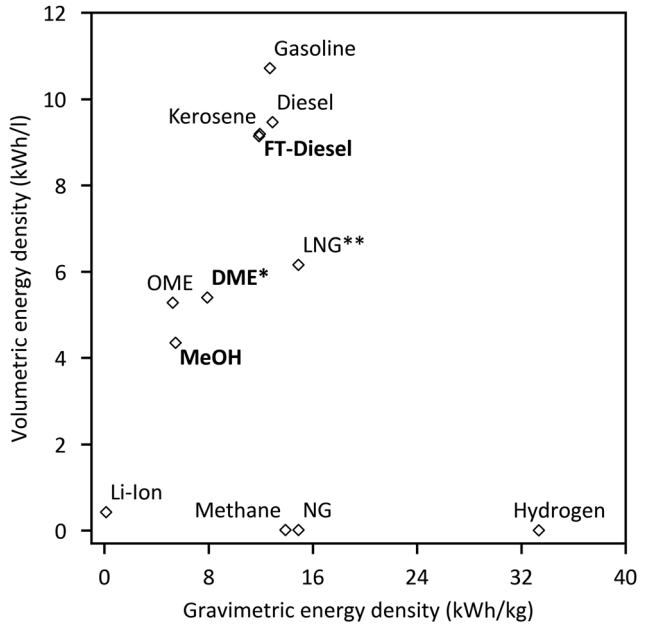

Fig. 1 Gravimetric and volumetric energy density of different fuels and Li-Ion batteries for comparison (*at 10 bar; **at −160 °C).[13–15]

regarding the interaction between the process steps is still scarce. While Section 1.1.1 will give a short introduction about water electrolysis, and Section 1.1.2 will summarise current carbon capture technologies, the article will focus on the synthesis and product upgrade within the PtL context. Three of the most prominent examples of PtL products will be discussed: methanol, dimethyl ether (DME) and Fischer–Tropsch-fuels (FT-fuels). For each, synthesis from CO and H_2 (syngas)† is an established technology with decades of operational experience.

However, using PtL as CCU technology raises the question of a direct synthesis from CO_2. There are several elaborate reviews and books on fuel production routes from CO_2 in general,[16–20] catalysts for the hydrogenation of CO_2 to methanol,[21–23] DME or FT-fuels, as well as conventional and innovative processes for the production of methanol,[24–29] DME[27,30–34] or FT-fuels[35–37] from natural gas, coal or biomass derived syngas. More recently reviews that evaluate the state of PtL with regard to the life cycle,[38] advances in catalysis and reaction routes,[39] as well as the area of application[40,41] have been published. However, a comprehensive technology review summarizing the status of synthesis technologies and evaluating new developments and research needs to address challenges within the context of PtL is missing.

Therefore, the review will provide an extensive overview of the synthesis for each product considering thermodynamics, catalysts, commercial converter designs and processes as well as the product upgrade. The goal is to provide boundary conditions for future concepts with regard to performances, catalyst life times and reactor and process designs. It should guide further developments and enable benchmarking. In addition, current research addressing the main challenges of the synthesis within the PtL process, *i.e.* CO_2 utilisation and flexible operation, will be discussed. An overview of existing PtL pilot plants is given and operation experiences are summarised. Moreover economic aspects and techno-economic analysis for medium-term commercialisation are discussed.

1.1. Power-to-liquid main process steps

1.1.1. Water electrolysis. A recent review of the current status of water electrolysis for flexible energy storage applications is given elsewhere by the authors.[42] The main water electrolysis technologies are alkaline electrolysis (AEL), proton exchange membrane electrolysis (PEMEL) and solid oxide electrolyser cell (SOEC). The major parameter regarding efficiency, investment costs, lifetime and flexibility are summarised in Table 1.

As can be seen, AEL and PEMEL are available on MW-scale and several flexible operated pilot plants have been realised in the last years. The hydrogen produced reaches very high purity of more than 99% (above 99.999% possible). Moreover pressurised electrolysers up to an operational pressure of 60–80 bar are commercially available reducing the need for compression.

SOEC on the other hand offers thermodynamic advantages due to operation at high temperatures theoretically enabling electric efficiencies above 100% based on LHV. However, industrial production is only in an early stage and pressurised operation is still in the research phase. In addition, SOEC can be used for syngas production *via* high temperature co-electrolysis of H_2O and CO_2. In this case hydrogen and carbon monoxide are formed. This variant is especially interesting for PtL processes using syngas and the application of co-electrolysis has already been discussed for PtSNG,[43] PtMeOH,[44] and PtFT-fuels.[45] However, degradation of the cell remains a challenge for commercialisation, making a better understanding of the reaction and degradation mechanism necessary

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† Within the context of this article synthesis gas, which mainly includes CO and H_2 will be referred to as 'syngas'.

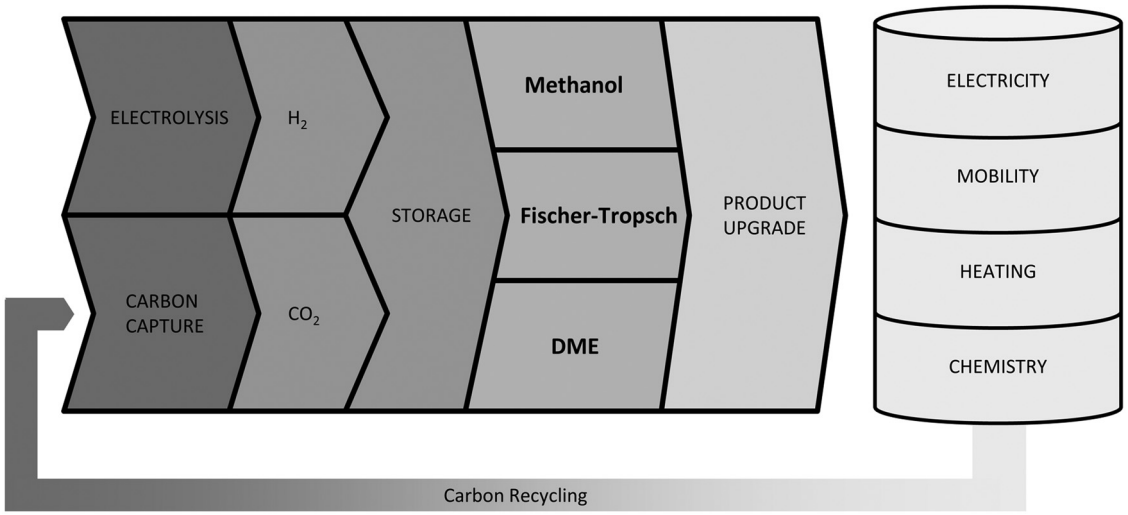


Fig. 2 Power-to-X pathways to transform electricity to chemicals *via* electrolysis and synthesis. The power-to-X principle connects different energy sectors *via* carbon capture from industrial plants or chemical/natural (photosynthesis) CO_2 scrubbing from the air, the carbon cycle can be closed so that no additional CO_2 is emitted. This paper regards paths to transform electricity to liquid fuels for the mobility sector.

Table 1 Summary of main parameters for current water electrolysis technologies summarised from Buttler and Spliethoff (2018)[42]

	AEL	PEMEL	SOEC
Operation parameters			
Cell temperature (°C)	60–90	50–80	700–900
Typical pressure (bar)	10–30	20–50	1–15
Current density (A cm^{-2})	0.25–0.45	1.0–2.0	0.3–1.0
Flexibility			
Load flexibility (% of nominal load)	20–100	0–100	−100/+100
Cold start-up time	1–2 h	5–10 min	Hours
Warm start-up	1–5 min	<10 s	15 min
Efficiency			
Nominal stack efficiency (LHV)	63–71%	60–68%	100%[a]
Nominal system[b] efficiency (LHV)	51–60%	46–60%	76–81%
Available capacity			
Max. nominal power per stack (MW)	6	2	<0.01
H_2 production per stack (N m^3 h^{-1})	1400	400	<10
Durability			
Life time (kh)	55–120	60–100	(8–20)
Efficiency degradation (% per a)	0.25–1.5	0.5–2.5	3–50
Economic parameters			
Investment cost (€ per kW)	800–1500	1400–2100	(>2000)[c]
Maintenance costs (% of investment costs per year)	2–3	3–5	

[a] Operating at thermoneutral voltage. [b] Including auxiliaries and heat supply (SOEC). [c] High uncertainty due to pre-commercial status of SOEC.

as well as further material developments. An overview of the technology, material and current developments is given by Zheng *et al.*[46]

1.1.2. Carbon capture. Carbon capture will not be discussed in detail within this review, however a series of comprehensive review articles have been published about carbon capture and storage in Energy & Environmental Science for additional information.[47–49] Table 2 gives a short summary of some prominent carbon capture technologies that can be applied in PtL processes.

In general CO_2 for the synthesis can either be captured from point sources with a high CO_2 partial pressure, *e.g.* power plant or industry exhaust gases, or can be obtained directly from the air *via* direct air capture (DAC) technologies. DAC requires much higher energy inputs and the processing of greater gas volumes.[49] Consequently, costs are expected to be up to ten times higher,[50] however, due to learning effects and process improvement cost competitiveness could be reached by 2040.[51]

It is difficult to estimate the costs of CO_2 from different carbon sources due to the variety of capture technologies,

Table 2 Examples of prominent carbon capture technologies suitable for PtL[47–49]

Methods	Description
Capture from point sources	
Absorption technologies	CO_2 is removed from flue gases by absorption in liquid solvents. Typically aqueous amine solutions are used and water-soluble salts are formed in the process. Afterwards CO_2 can be recovered in a desorption column.
Adsorption technologies	Solid adsorbents are used for the CO_2 capture. Typical adsorbents are carbons, alumina, silica and zeolites but also new adsorbents are investigated *e.g.* polymers. Cyclic processes are implemented for the loading and regeneration of the adsorbents in particular pressure vacuum swing adsorption and temperature swing adsorption. Furthermore, so-called carbonate looping processes, which are based on chemisorption can be used. The most prominent process is calcium looping (using CaO as sorbent). CO_2 is captured forming calcium carbonate. The sorbent is afterwards regenerated *via* calcination releasing the CO_2.
Direct air capture	
Wet air capture	CO_2 from ambient air is absorbed in a liquid solution within packed-columns, convection towers or spray-tower contractor systems. An example is the soda/lime process based on a sodium hydroxide solution.
Dry air capture	Typically solid organoamine based adsorbents are used for the CO_2 capture. Desorption occurs at elevated temperatures in an inert gas stream, however only a diluted CO_2 stream is generated. Newer developments focus on improving regeneration conditions.

Table 3 Estimated cost of CO_2 from different carbon sources based on Fasihi *et al.*,[51] Budinis *et al.*[52] Trost *et al.*[53] and Cormos *et al.*[54]

Carbon source	CO_2 cost in € per t_{CO_2}
Coal gasification power plant	28–40
Coal-fired power plant	31–49
Gas-fired power plant[a]	47–90
Refineries & NG processing[a]	18–71
Steel mill	70–73
Cement production	58–87
Biogas plant[b]	0–90
Direct air capture[c]	222–268

[a] In $€_{2015}$, exchange rate €/$ of 1.11. [b] Approx. costs for the seperation of CO_2 from biogas, currently an unused waste product. [c] Estimation for 8000 full load hours in 2020 and 10% learning curve effects.

location-specific circumstances and variable energy and utility costs. For a first indication Table 3 summarises recent cost estimates for CO_2 from different sources.

1.1.3. Fundamentals of synthesis loop and reactor design. The conversion of carbon monoxide, carbon dioxide and hydrogen to either methanol, DME or FT-fuels is strongly exothermic. Therefore, the equilibrium conversion is favoured by low temperatures while the reaction kinetics improve with higher temperatures. As a result, heat removal and recycling of unconverted syngas (which also dilutes the feed gas and reduces the temperature rise in the reactor)[55] is required to achieve a reasonable conversion.

A simplified flowsheet of a conventional synthesis loop applied in most cases is depicted in Fig. 3. The fresh feed gas is mixed with the recycle gas stream. This feed stream is optionally preheated before entering the synthesis reactor partially converting the syngas. The reacted gas is then cooled and liquids are separated by condensation. This crude product has to be further purified to meet the given specifications. The unreacted syngas is recycled with a small part being purged to avoid an enrichment of inert gases.

The optimal feed gas composition in terms of maximum conversion is usually defined by the stoichiometric number (SN) determined by the synthesis reactions. The stoichiometric

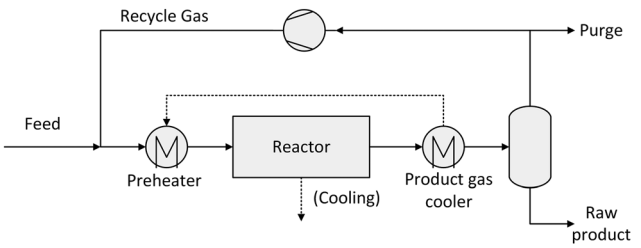


Fig. 3 Conventional synthesis loop in methanol, DME and FT-processes.

number is given by the ratio of the reactants hydrogen, carbon monoxide and carbon dioxide (in molar fraction):

$$SN = \frac{H_2 - CO_2}{CO + CO_2} \quad (1)$$

In general, similar reactor designs (and combinations of them) are applied for methanol, DME and Fischer–Tropsch synthesis. Table 4 gives an overview of typically used reactor designs. The various commercial reactor designs and the particular characteristics of methanol, DME and FT-fuels synthesis routes are discussed in the following sections in detail.

1.1.4. Alternative synthesis technologies. This review is focusing on thermochemical synthesis for fuel production, which also represents the commercial standard for the discussed products. Nonetheless, alternative synthesis technologies and processes have gained research interest in recent years and while they have not achieved a technology readiness for large-scale application yet, significant improvements have been made.

One alternative is the electrochemical conversion of CO_2 to value added products. CO_2 can be directly reduced to different products including lower alkanes, formic acid or methanol.[60] However, low solubility of CO_2 and CO species in aqueous electrolytes currently limits the formation of long-chain hydrocarbons.[61] Thus, of the products under consideration in this review, only methanol has been discussed for direct electrochemical production. Alternatively, a two-stage process is possible where CO_2 is converted to CO in a first electrolyser,

Table 4 Overview of reactor designs applied for methanol, DME and Fischer–Tropsch synthesis[24,35,56–59]

	Fixed bed reactor (FB)					
	Adiabatic FB	Quench	SRC[a]	GCC[b]	Fluidized bed reactor	Liquid-phase reactor
Operation mode	Adiabatic	Polytropic	Isothermal	Polytropic	Isothermal	Isothermal
State of catalyst	Catalyst bed	Catalyst bed divided into sections	Catalyst filled into shell or tube side	Catalyst filled into shell or tube side	Catalyst in fluidized bed	Catalyst suspended in liquid hydrocarbons
Cooling concept	Series of reactors with intercooling	Cold feed gas is injected between sections	Cooled by evaporation of water on opposing side of the catalyst	Cooled by preheating cold gas passing through the catalyst bed in tubes	Submerged coil heat exchanger	Liquid hydrocarbons as heat transfer medium; heat removed in internal heat exchanger
Advantages	Simple scale up, no mechanical stress of catalyst, defined residence time	No mechanical stress of catalyst, improved temperature control	Very efficient heat recovery, good heat transfer and temperature control		High heat transfer coefficients, uniform temperature distribution	Excellent heat transfer performance, isothermal reaction profile
Disadvantages	Lower heat transfer, danger of hotspots, several pressure vessels and piping	Large catalyst volume	Large number of tubes, expensive	Limited heat transfer	Non-uniform residence time (bubble formation), attrition of catalyst, erosion of internals, difficult scale up	High mechanical stress of catalyst

[a] SRC = steam raising converter. [b] GCC = gas-cooled converter.

which then reacts to other products in a second electrolyser under alkaline conditions.[61] Spurgeon *et al.*[45] compared the direct and two-stage electrochemical production of ethanol as well as the direct production of formic acid and a PtL process with an electrochemical reduction to CO with subsequent thermochemical FT-synthesis. Their results indicate that the PtFT-fuels process offers cost advantages over electrochemical conversion and that a two-stage electrochemical synthesis has slight cost benefits over a direct conversion. Current advances and remaining challenges for the direct electrochemical conversion as well as economic aspects have been reviewed by Birdja *et al.*[60] and Durst *et al.*[62] A review of the two-stage process is provided by Jouny *et al.*[61]

In comparison to electrochemical conversion, photocatalytic CO_2 conversion mimics natural photosynthesis by using solar energy to directly convert CO_2 without producing electricity first. The photocatalytic synthesis occurs in three main steps, the solar-light absorption, charge separation and migration followed by the catalytic CO_2 reduction.[63] Similar to electrochemical CO_2 conversion, the main products are CO, methane and methanol.[64] Long-chain hydrocarbons comparable to FT-fuels or DME have not been produced yet. The main advantage of photocatalytic CO_2-conversion is the independence from an external electricity source, allowing a higher flexibility in the choice of the location.[65] However, solar-to-fuel efficiency remains low and costs of the photocatalytic devices high (mainly due to noble metal-based catalysts).[66,67] A review on photocatalytic devices has been given by Jiang *et al.*[66] Wang *et al.* discuss the recent progress for different photocatalytic CO_2 conversion pathways[64] and reviews on recent catalyst developments have been published.[63,67]

In addition to the above-mentioned pathways of fuel synthesis there are biological processes which can be used for the production of biofuels. Hereby the use of bio membrane-reactors and biocatalysts (enzymes) can be differentiated.[68] For example, biological conversion of methane to methanol,[69,70] the synthesis of petroleum-like hydrocarbons[71] or biofuels from microalgae[72] could be achieved. Currently, there are several challenges that need to be addressed, such as dissolving CO_2 in or poisoning of the used solution.[70] In a recent review by Bhatia *et al.*, several biological ways of CO_2 capture and conversion as well as their associated challenges are discussed.[73]

1.2. Considered power-to-liquid fuels and comparison with conventional fuels

The properties of methanol, DME and FT-fuel (diesel fraction) are compared with conventional fuels in Table 5. Methanol can be used in both diesel and Otto motors. It is miscible with gasoline and has a higher octane number, which increases the efficiency of combustion. While the local emissions are lower in methanol combustion, the energy density is only about 50% of gasoline and methanol is disadvantageous in terms of corrosion. Furthermore, methanol has no lubricating effect on the motor. Methanol-gasoline blends with high methanol percentage (M85) as well as pure methanol (M100) have been successfully used in Otto motors with little modification for example in Brazil and Sweden. Methanol can also be used in fuel cells (direct or with reformer) with a very good efficiency.

Dimethyl ether is the simplest ether in nature. At ambient conditions, it is gaseous, just like LPG, but under moderate pressure (5 bar) it is in a liquid state. As it has no carbon bindings and contains no sulphur, soot and sulphur emissions are avoided. DME can be used in diesel engines, but modifications have to be made to the infrastructure to keep the fuel in a liquid state.[77] The energy density is lower, but the cetane number is higher than for conventional diesel, hence DME is

Table 5 Comparison of conventional and synthetic fuels

Properties	Gasoline	Methanol	Diesel	FT-fuel (diesel)	LPG	DME
Aggregate	Liquid	Liquid	Liquid	Liquid	Gaseous (liquid under pressure, 5–10 bar)	Gaseous (liquid under pressure, 5 bar)
Chem. formula	C_5–C_{12}	CH_3OH	C_{10}–C_{23}	C_{10}–C_{23}	C_3–C_4	CH_3OCH_3
Miscibility		In gasoline and diesel		In diesel		In LPG; in diesel
Degradable	No	Yes	No	Yes		Yes
Pollution		Oxygen content reduces local emissions	High soot and NO_X emissions	Less hydrocarbon, CO and particle emissions	NO_X emissions −80%, KW emissions −50% compared to gasoline	No C–C binding → almost no particle emissions
Density (g l^{-1})	715–780	791	815–855	770–860	540 (at 10 bar)	668
Boiling point at 1 atm (°C)	25–215	64.7	170–380	150–320	−42 to −0.5	−24.9
Vapour pressure at 20 °C (bar)	0.45–0.9	0.37	0.01–0.1	0.01–0.1	2.1–8.3	5.3
LHV (MJ l^{-1})	31.2–32.2	15.4–15.6	35.3–36	33.1–34.3	24.84	18.2–19.3
Octane number[74]	90–95	110–112	—	—	105–115	—
Cetane number	—	5 (low)	45–53	70–80	—	55–60
Ref.	75 and 76	75 and 76	75 and 76	74–76	75 and 76	75 and 76

seen as a promising option for heavy-load traffic. Further, DME might have advantages in terms of maximum energy efficiency compared to MeOH and FT-fuels.[78]

FT-based diesel has very similar properties compared to conventional diesel, except that it contains no sulphur and no aromatics, which decreases emissions drastically. Hydrocarbon, CO, NO_x and particle emissions are only 40–80% of those in conventional diesel combustion.[35] In the production process, the product can be adapted to the motor, which can increase the combustion efficiency. Due to the very similar properties, infrastructure and motors of conventional diesel can be used without adaption. FT-synthesis can be realised with biogenic feedstock or, to increase the output of conventional fuel production, it can be applied to stranded gas. Its products can be blended with low-quality conventional diesel to increase the yield of refineries.[74] The similarity with conventional diesel also causes some drawbacks. The product of FT-synthesis has to be fractionated to gain valuable products. Also, some products have to be treated further, *e.g.* in hydrocrackers. However, this also means that other products like diesel, gasoline and waxes can be produced on the same route.

2. Methanol synthesis

2.1. State of the art methanol synthesis

2.1.1. Theory. The industrial-scale synthesis of methanol began in 1923 at the facilities of BASF (Leuna Werke, Ludwigshafen/Germany). The process used a sulphur-resistant ZnO/Cr_2O_3 catalyst and operated at high pressures (250–350 bar) and temperatures (320–450 °C).[59,79] In the early stage of industrial production, methanol production from CO_2 has also been performed (1927, USA, Peoria, Commercial Solvents Corporation, 4000 t per a).[80] Since the 1960s, the high-pressure process was replaced by the now exclusively used low-pressure processes (50–100 bar) first commercialised by ICI (now Johnson Matthey).[55] The low-pressure process is operated above 200 °C determined by the activity of the applied copper catalysts and below 300 °C limited by its thermal stability. Earlier works, especially by Klier and coworkers (1982),[81] indicate CO as the major source of methanol synthesis. Based on isotope tracing experiments and kinetic observations today most scientists agree that methanol formation proceeds predominantly *via* CO_2 hydrogenation according to the following reaction:[26,82–93]

$$CO_2 + 3H_2 \rightleftharpoons CH_3OH + H_2O \quad \Delta H = -49.4 \text{ kJ mol}^{-1} \tag{2}$$

While CO_2 hydrogenation incorporates the formation of water, CO is converted to CO_2 *via* the reverse water gas shift reaction (RWGS) by consumption of water:

$$CO_2 + H_2 \rightleftharpoons CO + H_2O \quad \Delta H = 41.2 \text{ kJ mol}^{-1} \tag{3}$$

The CO conversion can be expressed as:

$$CO + 2H_2 \rightleftharpoons CH_3OH \quad \Delta H = -90.6 \text{ kJ mol}^{-1} \tag{4}$$

Both synthesis reactions are exothermal and involve a decrease of volume. Thus, methanol formation is favoured by low temperatures and elevated pressures. CO hydrogenation is significantly more exothermic than CO_2 hydrogenation resulting in a higher cooling demand. The maximum conversion is determined by the chemical equilibrium shown in Fig. 4 for a CO- and CO_2-based feed gas. As seen, the equilibrium limited methanol yield of CO_2 (18–58% at 200–250 °C, 50–100 bar) is substantially lower than that of CO (55–89%).

Based on equilibrium considerations, the highest methanol yield results from an educt gas composition with a stoichiometric number of 2 in the absence of CO_2, while for CO_2-based feed gas a H_2/CO_2 ratio of 3 to 1 is optimal. In practice, a slight excess of hydrogen given by an adjusted stoichiometric number of 2.02–2.1 is used, as this has been found to improve the space time yield of the catalyst and avoids by-product formation by a deficiency in hydrogen.[55,59,94,95] Additionally, experience

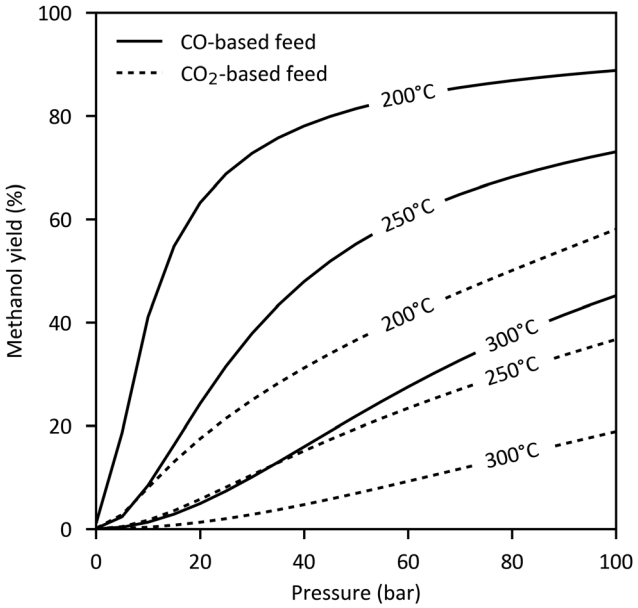


Fig. 4 Equilibrium conversion of CO- or CO_2-based feed gas to methanol.

with commercial copper-based catalysts reveal a maximum conversion at low CO_2 contents between 2–5%.[55,81,85,88,96] For higher CO_2 concentrations the methanol synthesis is inhibited by water formed in the reverse water gas shift reaction.[26,92]

2.1.2. Catalysts. Currently, all commercially applied low-pressure catalysts are based on CuO and ZnO in most cases on a carrier of Al_2O_3 with variable stabilising additives and promoters like Zr, Cr, Mg and rare earth metals. A detailed discussion of alternative catalyst formulations and promoters is given in several reviews.[20–22,24,26]

An overview of the composition, operating conditions and performance of some catalysts is listed in Table 6. The reported space time yields (STY) for CO syngases are in the order of 0.7–2.3 kg of methanol per litre of catalyst per hour (40–100 bar GHSV around 10 000 h^{-1}) which is reported from commercial plants as well.[97–99] CO_2-based feed gases usually show lower STY in the order of 0.4–0.8 kg l_{cat}^{-1} h^{-1}.

The main suppliers of conventional methanol catalysts are Johnson Matthey (Katalco 51-series), Clariant (MegaMax 700/800/NJ-1), Haldor Topsøe (MK 121 and MK 151-Fence) and Mitsubishi Gas Chemicals (M5-5, M6).[24] The catalysts are usually offered in a tablet shaped (*e.g.* 6 × 4 mm[100,101] or 5.1 × 5.3 mm[102]) and have a bulk density in the range of 1000–1300 kg m^{-3}.[98,102–106]

Modern copper catalysts achieve a high selectivity (apart from water) of over 99%[24,113] (up to 99.9% reported by Haldor Topsøe)[114] for standard syngas. Generally the low content of impurities is remarkable as the formation of typical by-products like higher alcohols (predominantly ethanol), esters (mainly methyl formate), ether (dimethyl ether), ketones (mainly acetone) and hydrocarbons are thermodynamically favoured over methanol (except for formaldehyde and formic acid).[26] By-product formation is promoted by catalyst impurities (residual amounts of alkalis), high pressures, high temperatures (requiring adequate control of the converter temperature), higher CO/H_2 and CO/CO_2 ratios, as well as lower space velocities (higher residence time).[115,116]

In industrial applications the common catalyst life time is 4–6 years[95,117,118] with up to 8 years being reported.[102] Life time is limited by catalyst deactivation caused by poisoning and thermal sintering. Major poisons for copper catalysts are sulphur compounds (blocking of active sites) and chlorides (acceleration of sintering). The tolerance against sulphur is higher than chlorides, as sulphur is scavenged by the ZnO component of the common Cu/ZnO catalysts.[119] Other poisons reported in industrial applications are, for example, arsenic or carbonyls (formed at high CO concentrations at inappropriate types of steel)[120] resulting in a decrease of selectivity due to promotion of Fischer–Tropsch side reactions.[121] Typical gas

Table 6 Composition and performance of some commercial methanol catalysts

Licensor	Composition[a] CuO/ZnO/Al_2O_3 (wt%)	T (°C)	p (bar)	GHSV[b] (×10^3 h^{-1})	STY (kg l_{cat}^{-1} h^{-1})	Byprod.[c] (wt%)	Ref.
CO-based feed gas[d]							
Süd Chemie	65–68/21–23/10–12	250	50	10	1.03–1.10	0.3–0.4	104
BASF	40/52/9	230	50	10	1.43	<0.6	106
ICI	62/32/6	240	51	9.6	0.43	<0.2	107
Shell	65/29/0/6 (di)	250	53	11.5	0.69	Trace	108
		300	53	10.9	1.01	<1.3	
Casale	30/50/3/16 (Cr_2O_3)	250	100	12.5	1	n.a.	109
MGC	34/26/3/36 (Zr)	264	100	10	2.22	n.a.	103
Topsoe (MK-121)	>55/21–55/8–10	200–310	39–122	n.a.	n.a.	n.a.	100
Topsoe (MK-101)	×/×/×	221[e]	67[e]	n.a.	1.02	n.a.	105
ICI	40/41/0/19 (Cr_2O_3)	260	41	6–7	0.26	<0.1	110
		260	41	10	0.38	<0.1	
		260	81	10	0.77	<0.1	
Metallgesellschaft	60/30/0/10 (Cr_2O_3)	250	41	9.8	1.42[f]	n.a.	111
		250	101	9.8	2.28[g]	n.a.	
NIRE/RITE	×/×/×/ZrO_2/SiO_2	250	50	10	1.7	n.a.	112

[a] Partially calculated based on given atomic ratio. [b] Gas hourly space velocity as ratio of feed gas volume flow under standard conditions to catalyst volume. [c] Except water. [d] $H_2:CO > 2:1$ (H_2 65–90%, CO 5–25%, CO_2 4–14%). [e] Applicable temperature range 200–310 °C, pressure range 20–150 bar. [f] Activity at start of run, decreased to 1.28 kg l^{-1} h^{-1} after 1000 h. [g] Activity at start of run, decreased to 2.20 kg l^{-1} h^{-1} after 1000 h.

Table 7 Gas purity requirements for Cu/ZnO methanol catalyst

Component	Purity requirement	Ref.
H_2S	<0.05–0.5 ppm	24, 105, 113, 119 and 122
HCl	1 ppb	113
Metal carbonyl	Few ppb	119
Particles	<0.1 mg N m^{-3}	123
Tar	<1.0 mg N m^{-3}	123
Alkalis	<0.25 mg N m^{-3}	123

purity requirements are summarised in Table 7. These impurities are commonly removed by prior gas cleaning. Additionally, guard beds (*e.g.* ZnO targeting sulphur) are often installed to protect the catalyst.[113,120,122] As a result, in normal operation of methanol synthesis plants deactivation is dominated by thermal sintering.[113] A growth of copper crystallite size is reported at temperatures higher than 227 °C and above 300 °C a sintering of ZnO is also observed.[119]

Typical curves of deactivation of commercial methanol catalysts are illustrated in Fig. 5. After an initial strong decline in catalyst activity during the first 1000 h (−14 to −21%/1000 h), the deactivation slows down, resulting in a mean deactivation rate of about 2%/1000 h over 3 years (based on presented data). Operation with CO_2-based feed gas indicates comparable stability to syngas, with a decrease in catalyst activity of 8% in the first 400 h,[124,125] but there is still a lack of long-term studies of CO_2-based pilot plants.[126] Intensive research on catalysts results in enhanced activity and stability. Haldor Topsøe reports a 20 percentage-point increase in activity at the end of life in each case from the MK-101 to the MK-121 to their latest methanol catalyst MK-151 Fence.[127] To maintain a constant production of methanol in practice, the operating temperature (or the pressure) is gradually increased during time on stream.[105,121,128]

2.1.3. Commercial converter designs and processes. Today's methanol converter development is driven by the trend of increasing capacities to make use of economy of scale effects. So-called Mega, Giga or Jumbo methanol plants of 5000 t per d to 10 000 t per d have been developed. Compared to a 2000 t per d plant the specific costs can be reduced by 30% for a 5000 t per d plant and by 40% for a 10 000 t per d plant.[131] Hereby the maximum capacity of the methanol converter is limited energetically by the pressure drop as well as by manufacturing and transportation constraints of the pressure vessel. Methanol production is mainly based on syngas from reforming natural gas. In the last few years methanol production *via* coal gasification has been growing too, driven by China.

The main challenge in reactor design is to economically remove the heat of reaction avoiding by-product formation and achieving a high conversion rate by low outlet temperatures and a good energy efficiency by internal heat recovery.[26]

A summary of the main commercial reactor types (Johnson Matthey Davy Technologies is missing due to scarce information)[132] is given in Table 7. Respective simplified reactor schemes are illustrated in Fig. 6 and detailed temperature profiles are shown in Fig. 7.

Current reactor designs are dominated by quasi-isothermal steam-raising fixed bed reactors (SRC) which replaced the former design of quench reactors due to the higher catalyst volume utilisation, improved heat recovery and temperature control. SRC were introduced for methanol synthesis for the first time by Lurgi (now part of Air Liquide) at the beginning of the 1970s.[133] The Lurgi design is based on a tubular reactor (Fig. 6a). The feed gas flows in an axial direction through the tubes filled with catalyst cooled by the surrounding boiling water on the shell-side. About 80% of the reaction heat is converted to medium pressure steam.[99] Similar designs are currently offered by Johnsson Matthey Davy Technologies[132] and Haldor Topsøe.[114] Haldor Topsøe optionally proposes an adiabatic catalyst bed installed upstream in a separate vessel or on top of the upper tube sheet resulting in an optimised temperature profile.[134] The maximum single-train capacity of SRCs is limited to 1500–2200 t per d (the range mainly depends on the syngas activity).[114,117,135] Further quasi-isothermal fixed bed reactor concepts are licensed by Linde, Toyo, Methanol Casale and Mitsubishi Heavy Industries as discussed in the following.

The Linde Variobar process is based on their coil-wound heat exchanger with catalyst loaded on the shell-side and boiling water circulating through the tubes (Fig. 6b).[136] Linde claims that this design features the highest catalyst per reactor volume ratio of all isothermal reactors and high heat transfer coefficients due to the cross flow reducing the heat transfer area.[137]

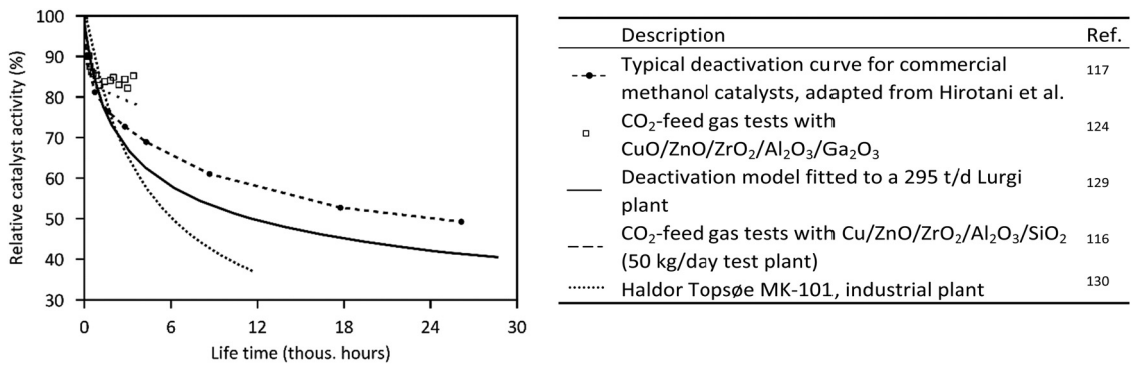


Fig. 5 Deactivation curves of commercial methanol catalysts.[116,117,124,129,130]

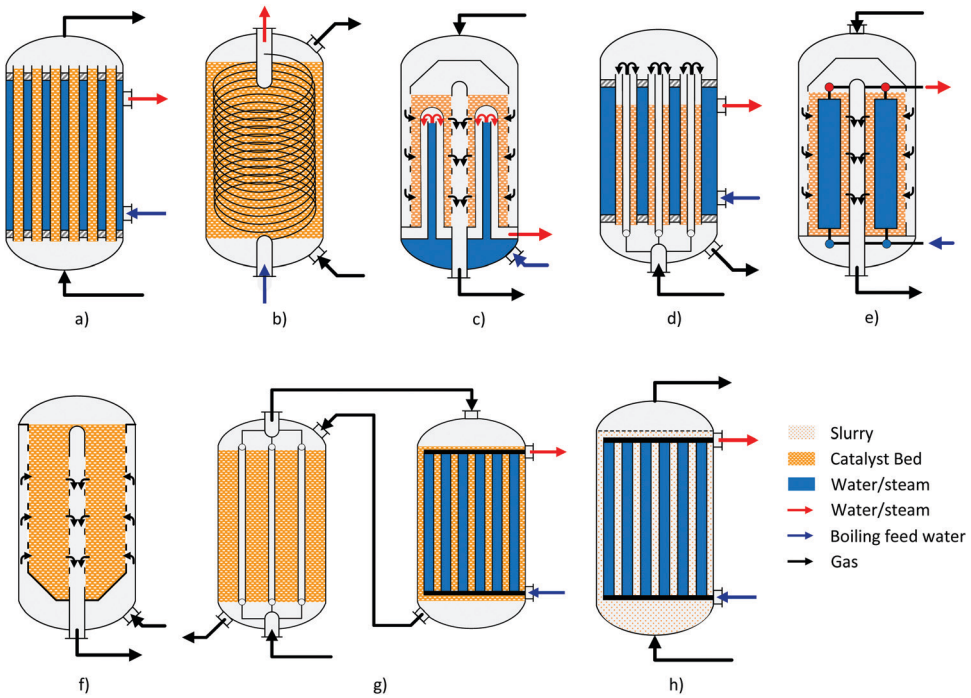


Fig. 6 Simplified reactor layout of (a) Lurgi tubular reactor, (b) Linde Variobar, (c) Toyo MRF, (d) Mitsubishi Superconverter, (e) Methanol Casale IMC, (f) Haldor Topsøe adiabatic reactor, (g) Lurgi MegaMethanol and (h) Air Products LPMEOH, adapted from Buttler 2018.[138]

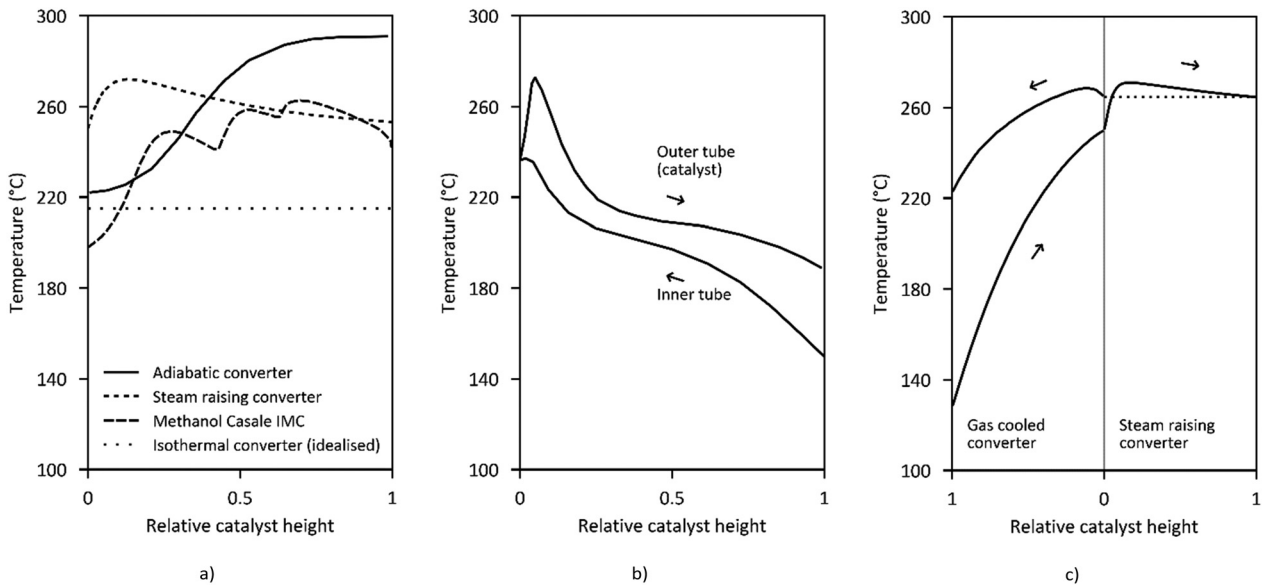


Fig. 7 Temperature profiles along the reactor length of (a) adiabatic converter,[105] steam raising converter,[126] Methanol Casale IMC,[142] and isothermal converter (idealized), (b) Mitsubishi Superconverter,[143] (c) Lurgi MegaMethanol.[99]

Toyo Engineering Corporation (TEC) developed a Multi-stage Radial Flow (MRF) reactor jointly with Mitsui Toatso Chemicals (MTC).[139] It is characterised by the radial flow of syngas across the catalyst bed loaded in the shell-side, which is indirectly cooled by bayonet boiler tubes (Fig. 6c). Thermal stress is avoided by using duplex tubes with the coolant first flowing up the inner tube before being evaporated in the outer tube. The arrangement of the tubes results in concentric cooling zones. This allows the temperature profile to be adjusted close to the path of the maximum reaction rate curve.[117] The radial flow arrangement results in a very low pressure drop across the reactor (<0.5 bar).[117,139] This has the advantage of a simple scale-up by vertical extension of the reactor with single-train capacities of 5000 t per d.[117]

Mitsubishi Gas Chemical and Mitsubishi Heavy Industries have jointly developed the Mitsubishi Methanol Process based

on the so-called Superconverter (Fig. 6d). This complex reactor design combines the features of a tube gas cooling reactor and a steam-raising tubular reactor by using double-pipe tubes in a boiling water vessel (or even triple pipes).[140] The feed gas first flows upwards through the inner tube before passing downwards through the outer tube which is filled with catalyst.[141] Preheating the feed gas in the inner tube results in a specific temperature profile (Fig. 7b) along the catalyst bed with a high temperature (maximum 250–260 °C) near the inlet that gradually declines towards the outlet (240–250 °C) closely following the maximum reaction rate line.[140]

The Methanol Casale (pseudo) Isothermal Methanol Converter (IMC) is cooled by feed gas, boiling water or a combination of both inside hollow plates which are immersed in the catalyst bed (Fig. 6e). Independent temperature control of different parts of the plates is possible by adjusting the cooling fluid flow at different heights.[144] Thereby the quasi-isothermal temperature profile (Fig. 7a.3) can be modified to fit the maximum reaction rate curve.[128,145] For capacities up to 2000 t per d axial flow of the process gas is preferred.[146] For larger capacities up to 7000–10 000 t per d in a single converter, axial-radial flow configuration is used due to the lower pressure drop.[146–148]

Another concept which is especially suited for large capacities of 10 000 t per d in a single line are multi-stage adiabatic fixed bed reactors with intercooling as offered by Haldor Topsøe (Fig. 6f) and Kellogg (now part of KBR).[26] Very large capacities based on axial flow steam-raising converters are either realised by several reactors in parallel or by dual stage concepts like the MegaMethanol concept of Lurgi (Fig. 6g), with a capacity of 5000 t per d.[94] The syngas is partially converted to methanol in a steam raising reactor before it is further converted in a tubular reactor cooled in counter-current (or co-current) flow with the cold feed gas for the first reactor.[149] This has the advantage that only a small feed gas preheater (up to about 130 °C) is required which results in a decreasing temperature along the reaction path maintaining the equilibrium driving force for methanol production (Fig. 7c).

Innovative reactor concepts. Beside conventional fixed bed reactor design, alternative concepts have been investigated which are very well discussed by Hansen and Nielsen[26] and Riaz *et al.*[28] Examples are membrane reactors or sorption-enhanced reactor concepts[150,151] to overcome the equilibrium conversion by removal of products or fluidized bed reactors (*e.g.* 10 t per d demo plant[152]), trickle bed reactors[153] and slurry reactors,[121] which eliminate the diffusion limitations and offer a good heat transfer. To date, none of the alternative concepts have any industrial relevance. Membrane reactors modules are still in the fundamental research stage and suffer from low conversion (experimentally below 9%)[154–162] as well as durability and cost issues, while fluidized bed reactors have problems with attrition of the catalyst.[26]

The only one of these alternative concepts that has been tested on an industrial scale is the Liquid Phase Methanol (LPMeOH) process developed by Air Products and Chemicals (1997–2002, Eastman Chemical's Coal Gasification Complex, 235 t per d, documented in ref. 121 and 163–165). It is based on a slurry bubble column reactor with fine catalyst particles suspended in an inert mineral oil acting as an efficient heat transfer medium (Fig. 6h). The heat removal *via* an internal heat exchanger results in an almost isothermal operation of the reactor. The superior temperature control compared to conventional fixed bed converters enables the use of CO-rich syngases (in excess of 50% tested) without damaging the catalyst by excessive temperature peaks. However, the LPMeOH process has a poor reactor volume utilisation due to the limited catalyst loading (slurry concentration of 25–50 wt%).[166]

Haldor Topsøe demonstrated a novel methanol fixed bed reactor synthesis as part of the small-scale BioDME pilot plant in Pitea commissioned in 2010.[167,168] Once-through operation is achieved by a second radial converter after the conventional steam raising converter. The second converter is operated under methanol condensing conditions.[169] This results in a once-through methanol yield higher than 95% (130 bar) as the equilibrium constraint can be overcome.[170] The first experiments of this technology were already published in 1991.[171]

Operating conditions and performance of conventional converters. The operating conditions are mainly determined by the applied catalysts while the temperature profile along the reaction path and the heat recovery is determined by the reactor design (a good overview of characteristic conversion profiles of different converter types is given by Hirotani *et al.*).[117] Reported exemplary operating parameters and performance of converter designs are summarised in Table 8.

High operating temperatures improve the reaction kinetics (at the risk of by-product formation) while low temperatures at the reactor outlet are favoured by the equilibrium, maximising the overall conversion. Detailed temperature profiles of exemplary converters are given in Fig. 7. As seen, the Superconverter design as well as the MegaMethanol process achieve relatively low temperatures at the outlet and only a small preheater is necessary due to the internal preheating of the feed gas. The heat transfer coefficient and the peak catalyst temperature determine the steam pressure of the cooled reactors, which is typically below 300 °C. Resulting steam pressures are in the range of 16–60 bar (T_{sat} ~ 200–275 °C). Steam raising converters produce about 0.7–1.4 t of medium pressure steam per tonne of methanol.[55,121,143,148]

In order to achieve an almost complete carbon conversion rate (typically 93–98%, dependent on purge gas losses)[114,148,172] the rate of recycled unconverted gas to fresh feed gas must be between 3 and 5 (Table 8). The recycle gas ratio can be reduced in case of the two-stage MegaMethanol process to 2–2.7.[24,173] By condensation of methanol after the first reactor, the driving force of the methanol reaction is increased, allowing for a reduced size of the gas cooled reactor and a lower recycle gas ratio of about 1.7.[174] The recycle gas ratio affects the size of the equipment as well as the power requirement of the recycle compressor. Beside the reactor design (and catalyst activity), the recycle gas ratio is also dependent on the reactivity of the gas. Reported conversion per pass lie in the range of 20–50%

Table 8 Overview of main commercial methanol reactor technologies and exemplary operation parameters from literature

	Lurgi	Variobar	MRF-Z[a]	Super-converter	IMC[b]	Haldor Topsøe	Mega Methanol	LPMeOH[c]
Licensor	Lurgi	Linde	Toyo (TEC)	MGC and MHI	Methanol Casale	Haldor Topsøe	Lurgi	Air products and chemicals
Reactor type[d]	SRC	SRC	SRC	GCC/SRC	SRC	Adiabatic[e]	GCC and SRC	Slurry
Catalyst location	Tube-side	Shell-side	Shell-side	Double-pipes	Shell-side	Fixed bed	Shell-side	Shell-side[f]
Heat exchanger	Tubular	Tubular (spiral)	Bayonet tubes	Tubular	Plate	Intercooler	Tubular	Tubular
Flow	Axial	Radial	Radial	Axial	Axial/radial[g]	Radial	Axial	Axial
Stages	1	1	1	1	1	2–4	2	1
T^h (°C)	255/270	n.a.	240/280	190/270	225/280	290/n.a.	220/270	215[i]
p (bar)	50–100	50–150[j]	80–100	55–100	65–80[k]	50–100	75	30–50
Recycle ratio[l]	3–4	n.a.	n.a.	2–3	3	3–5	2–2.7	1–5
Pressure loss[m] (bar)	3[n] (295 t per d)	n.a.	0.3–0.5[o] (300–2500 t per d)	2.4–7.5 (10 t per d)	1.1 (axial[p]), 0.3 (radial[q])	n.a.	n.a.	3–4.8 (235 t per d)
Per pass conversion[r] (%)	36	n.a.	60	55–67	n.a.	n.a.	>80	20–50
Y_{MeOH}[s] (mol%)	6–7	n.a.	10	10–15	10.1–13.3[t]	7	11	8–12
p_{Steam} (bar)	29–43	40	n.a.	19–45	25–32	n.a.	50–60	16–25
Max. capacity (t per d)	1500–2200	4000	5000	n.a.	7000–10 000	10 000	5000–10 000[u]	Low
Industrial references	>55 plants licensed	8 plants	6 projects, 315–3000 t per d	9 plants, 520–5000 t per d	9 plants, 1350–3000 t per d[v]	>40 plants	>10 plants since 2001[w]	235 t per d demo-plant
Ref.	55, 98, 99, 114, 117, 126, 131, 135 and 177	24, 29, 118, 137 and 182	117, 139 and 181	97, 140, 143 and 178–180	95, 120, 128 and 146–148	55, 105, 122, 127 and 134	99, 173, 176 and 177	121

[a] Multi-stage indirect cooling and radial flow. [b] Isothermal methanol converter. [c] Liquid phase methanol. [d] Simplified reactor schemes are presented in Fig. 6. [e] Haldor Topsøe also offers steam raising converters. [f] Catalyst powder suspended in mineral oil. [g] <2000 t per d axial flow, >2000 t per d axial–radial configuration.[146] [h] Outlet/Peak catalyst temperature, temperature profiles see Fig. 7. [i] Due to catalyst aging. The demonstration unit was tested from 214–259 °C.[121] [j] <1000 t per d, p = 50–80 bar, >1000 t per d, p = 70–150 bar.[137] [k] Revamping projects with operating pressures of 46–100 bar[128,148] realised. [l] Defined as ratio of recycle gas to fresh make-up gas based on volume flow. [m] Reactor only. [n] Pressure loss of the overall synthesis loop is 3.5–4 bar.[55] [o] Pressure loss of the overall synthesis loop is 3 bar.[139] [p] 2000 t per d. [q] 3000 t per d. [r] Per pass conversion = $1 - (CO_{out} + CO_{2,out})/(CO_{in} + CO_{2,in})$ based on given molar flow rates (partially own calculations based on given inlet and outlet compositions). [s] Methanol concentration at reactor outlet. [t] Depending on the operating pressure (65–79 bar).[120] [u] >5000 t per d by two SRC in parallel and condensation of methanol before one GCC.[24] [v] Data from 2013.[142] In total 23 new projects and 22 revamping projects since 1993.[175] [w] First MegaMethanol plant with 5000 t per d commissioned in 2004.

for the LPMeOH-Process,[121] below 40% for conventional SRC (own calculation based on[98]), 55–67% for advanced SRC like Toyo MRF and Mitsubishi Superconverter (own calculations based on ref. 97 and 117) and above 80% for the two-stage MegaMethanol process.[94] The resulting methanol concentrations at the reactor outlet are in the range of 6–14 mol%.

Due to the energy penalty of compressing the recycle gas, the pressure loss across the reactor (and the preheater and product cooler) is important. Typical pressure losses are below 0.5 bar for radial flow converters (overall pressure loss of 3 bar for a 2500 t per d synthesis loop[139]) or between 2.4 and 7.5 bar for axial flow converters (Table 8).

The amount of purge gas is dependent on the inert gas fraction (N_2 and CH_4) in the feed gas. In the LPMeOH experiments about 2–6% of the recycle gas were purged (corresponding to 1–5% of the reactor inlet gas or 5–20% of the feed gas) keeping the fraction of N_2 and CH_4 below 5% in the reactor inlet gas.[121]

Selected gas hourly space velocities (GHSV), the ratio of feed gas volume flow under standard conditions to catalyst volume, are in the range of 6000–12 000 h^{-1} for single-stage steam raising reactors.[55,97,98,173] For two-stage concepts, increased GHSV of 14 000–24 000 h^{-1} are reported.[173] The LPMeOH demonstration plant was operated at lower space velocities in the range of 2000–5000 h^{-1} (own calculations based on ref. 121).

These operation experience of CO-syngas converters represent the basis for CO_2 conversion pilot plants as discussed in the next section.

2.1.4. Product upgrade. In order to meet the desired methanol purity specifications (fuel grade or chemical grade A/AA)[59] impurities have to be removed from the raw product by distillation. Raw methanol withdrawn from the synthesis loop is first flashed at a low pressure of 5–10 bar to drive out dissolved gases.[55] In case of fuel-grade methanol, a single column is sufficient to meet the specifications, which reduces the investment costs.[139] For chemical-grade methanol two or three column distillation (see Fig. 8) is applied.[24] In a pre-run column low boiling by-products are removed overhead and are used as a fuel gas together with the expansion gas. Methanol, water and high boilers are discharged from the bottoms and fed to the refining column where pure methanol is withdrawn at the column top.[55] The refining column can be split up into a pressurised column and an atmospheric column. Heat integration is possible by using the latent heat of condensation in the pressurised column condenser to heat the atmospheric column reboiler. The three column arrangement reduces the heating demand for the refining column by about 30–40% but implies

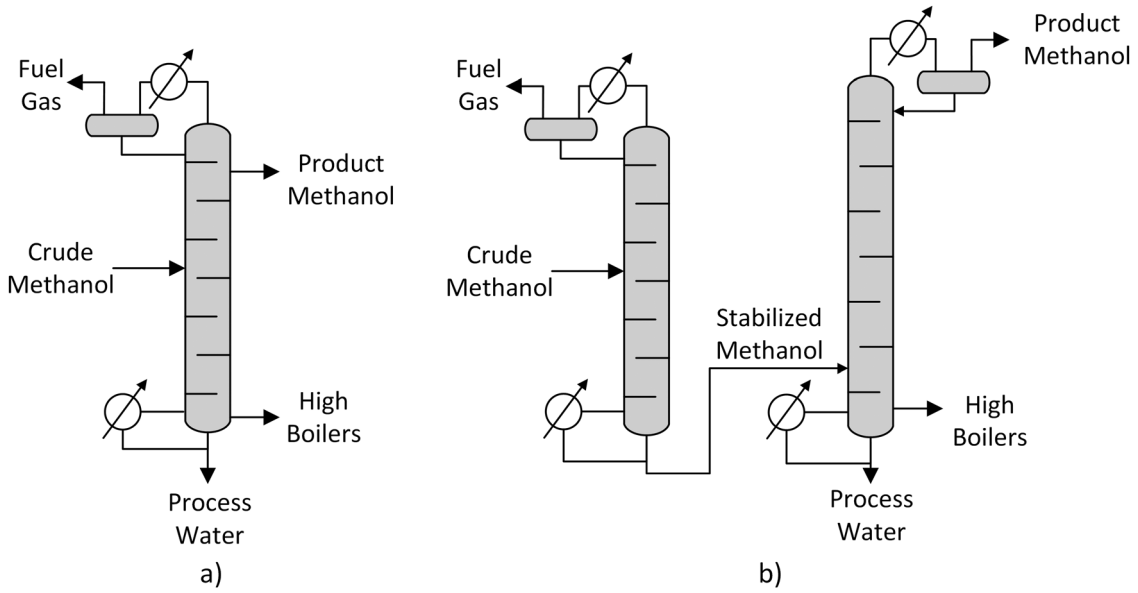


Fig. 8 Single column (a) and two-column (b) methanol distillation, own illustration based on Supp.[55]

higher investment costs.[24,139] A detailed description is given by Supp.[55]

2.2. Current research on power-to-methanol

2.2.1. Methanol from CO_2. Since the 1990s, catalysts and processes have been developed for large-scale conversion of CO_2 to methanol. For CO_2-based feeds standard methanol catalysts have been used,[126,183] however, new catalysts are under investigation which are adapted to the specific features of CO_2 conversion, especially the formation of water. Examples are multicomponent catalysts for CO_2-based feed gases.[92,112,124,184–186] Recent advances in heterogeneous catalysis of CO_2-based feed gases to methanol have been discussed in detail by several authors.[187,188] Table 9 gives an overview of some commercial catalysts with CO_2-based syngas.

Typical compositions of crude methanol produced from syngas and CO_2-based feed gas are compared in Table 10. The CO_2 conversion in reaction (2) is accompanied by formation of water resulting in significantly higher water contents of 30–40% of the crude methanol compared to 10–15% for syngas.[126,191] However, the formation of other by-products (see Tables 6 and 9) is lower for CO_2-based feed gas with contents below 0.05 wt% in crude methanol compared to 0.1–0.6 wt% for CO-syngas.

The higher selectivity of CO_2 conversion compared to CO conversion can be partly explained by lower peak temperatures in the reactor at the same jacket temperature as CO_2 hydrogenation is less exothermic. However, even at comparable peak temperatures by-product formation from syngas is about three times higher than for CO_2-based feed gas. Based on this, Pontzen *et al.* concluded that CO_2 conversion can be run at higher temperatures by less intense cooling in order to increase productivity.[126] Further, Göhna and König from Lurgi concluded that the energy demand of distillation can be reduced by about 20%.[191]

An overview of main operating and performance parameters of CO_2 to methanol pilot plants is given in Table 11. Operating parameters are in the range of conventional syngas processes ($T \sim 250$ °C, p = 50–100 bar, GHSV $\sim$ 10 000 h^{-1}). However, the reported carbon conversion rate is lower in several pilots ranging from 40–97% (with recycle). Most of the pilot plants are on a laboratory or small demonstration scale (<30 t per a). The only commercial plant has been commissioned by Carbon Recycling International in 2012 with a capacity of 4000 t per a.[193]

Table 9 Composition and performance of some commercial methanol catalysts with CO_2-based feed gas

Licensor	Composition $CuO/ZnO/Al_2O_3$ (wt%)	T (°C)	p (bar)	GHSV[b] ($\times 10^3$ h^{-1})	STY (kg l_{cat}^{-1} h^{-1})	Byprod.[c] (wt%)	Ref.
Süd Chemie	n.a.	250	80	10.5	0.6	0.04	126
NIRE/RITE[a]	$x/x/x/ZrO_2/Ga_2O_3$	250	50	18	0.75[b]	0.04	124
NIRE/RITE[a]	$x/x/x/ZrO_2/SiO_2$	250	30–70	10	0.4–0.8	0.05	112
NIRE/RITE[a]	$45/27/5/23(ZrO_2)/0.6$ (SiO_2)	250	50	10	0.76[c]	n.a.	186
JM Katalco 51-8	×/×/×	240	69–97	3.3–8.3	0.06–0.24	<0.05[d]	183
Süd Chemie (C79-05-GL)	×/×/×	260	80	8.1	0.66	n.a.	189
Commercial[a]	60/30/10	250	30	7.9[e]	0.6	n.a.	190

[a] H_2/CO_2 of 3 : 1(H_2 = 75%, CO_2 = 25%) with 22% CO_2 and 3% CO, $H_2/(CO + CO_2)$ = 2.4 and $CO_2/(CO + CO_2)$ = 2.4 after RWGS reactor (H_2 = 70.6%, CO = 17.6%, CO_2 = 11.8%). [b] Activity at start of run, decreased to 0.63 kg l^{-1} h^{-1} after 1000 h. [c] Activity at start of run, decreased to 0.70 kg l^{-1} h^{-1} after 1000 h. [d] In volume percentage. [e] Calculated based on given catalyst volume and reactor inlet flow.

Table 10 Raw methanol compositions of CO and CO_2-based feed gases[191,192]

	CO-Syngas	CO_2-based feed gas ($H_2:CO_2$ = 3)					
Temperature (°C)	250	250	255	260	230	250	270
Main components (wt%)							
Methanol	84.5	63.7	63.7	63.7	63.3	63.0	63.4
Water	15.4	36.2	36.3	36.3	36.7	36.9	35.6
Impurities (wt ppm)							
n-Paraffins	78	0	0	0			
Higher alcohols	626	89	105	148	28	45	92
Esters[a]	582	145	140	129	450	290	270
Ketones	24	0	0	0			
Dimethylether	61	14	18	24			
Total impurities (wt ppm)	1371	248	263	301	478	335	362
Methanol selectivity, except water (%)	99.84	99.96	99.96	99.95	99.92	99.95	99.94

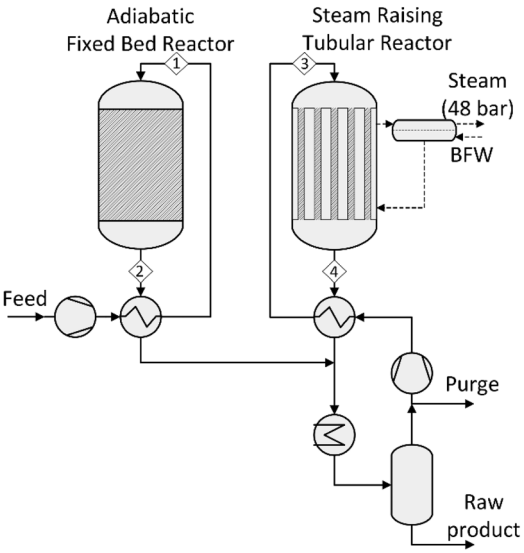
[a] Mainly methyl formate.

Table 11 Bench, pilot and demo plants for CO_2 to methanol processes

Institute/company, country, year	Capacity (kg d^{-1})	Reactor volume (l)	T (°C)	p (bar)	GHSV ($\times 10^3$ h^{-1})	Carbon efficiency[a] (%)	Recycle (ratio[b])	Description[c]	Ref.
Lurgi AG Germany, 1994[d]/2010	n.a.	n.a.	250[e]	80	10.5	94.0–96.5[f]	Yes (4.5)	Heated water jacket	126
NITE, RITE Japan, 1996/1998	0.9[g]	0.05[h]	250 (200–275)	50	18	n.a.	Yes	Reactor immersed in a sand bath	124 and 185
	50	4.6[i]	250 (230–270)	50 (30–70)	10 (5–20)	n.a.[j]	Yes[k]	Oil cooled reactor	112, 116, 184 and 192
Centre for Solar Energy and Hydrogen Research (ZSW) Germany, 1996	6.1[g]	0.4[h]	260	80	8.1	23	No	Electrical heated jacket	189 and 199
Korea Institute of Science and Technology (KIST) Korea, 1999/2004	75[l]	8.1	250–300[n]	51/61	n.a.	66.9–70.5	Yes	Prior RWGS reactor (electrical heated) and four SRC in parallel	196
	n.a.	n.a.	250	30	7.9[m]	53[n]	Yes (1[o])	Minipilot-plant	190
Mitsui Chemicals Inc. Japan, 2009	274[p]	n.a.	250	50	10	72–88[q]	Yes (2.6–3.2[r])	Pilot plant	197, 198 and 200
Northern Arizona University (NAU) USA, 2009/2014	<0.5[g]	0.05–0.08[h]	240	69–97	3.3–8.3	2.6–14.3	No	Electrical heated jacket	183
	1.6	0.15[s]	240(/260)	90	1.0 (0.6–12.0)[t]	40	Yes	Mobile test rig, electrolysis, electrical heated reactor jacket	201–203
Silicon Fire AG Switzerland, 2010	40	n.a.	265	80	n.a.	n.a.	n.a.	Pilot plant	204
Carbon Recycling International (CRI) Island, 2012	12 000[u]	n.a.	250	100	n.a.	n.a.	Yes	Commercial plant converting CO_2 from geothermal flue gas	205 and 206
CRI, MHI Germany, 2019	1000	n.a.	n.a.	n.a.	n.a.	n.a.	n.a.	Integrated in a coal power plant	207 and 208

[a] Carbon efficiency = $CH_3OH_{Product}/(CO + CO_2)_{MUG}$. [b] Recycle gas ratio = Recycle (mol s^{-1})/MUG (mol s^{-1}). [c] All plants based on fixed bed reactor technologies. [d] First CO_2 results by Göhna and König[191] in 1994. [e] Peak temperature in catalyst bed of 260–264 °C. [f] Per pass conversion of 35–45%. [g] Calculated based on catalyst volume and STY. [h] Catalyst volume. [i] d = 38.4 mm, l = 4 m; given catalyst volume = 0.5–3 l. [j] Per pass conversion of 13.6–16.7%. [k] Purge ratio = 0.5–1% of reactor inlet flow. [l] Designed for 100 kg d^{-1}. [m] Calculated based on cat. volume and inlet flow. [n] Calculation based on given material balance of ref. 190 who states a carbon efficiency of 89% not taking into account purge gas losses. Per pass conversion of 30%. [o] Due to prior RWGS, Purge ratio = 40% of MUG or 22% of reactor inlet flow. [p] 100 t per a. [q] Including recycling of the combusted purge gas. [r] Purge = 27–37% of MUG, 7–10% of reactor inlet flow. [s] Based on given dimensions (l = 30.5 cm, d = 2.5 cm). [t] Based on flow (1.5–30 NLPM) and reactor volume. [u] Plant opening with 1300 t per year, expanded in 2014 to 4000 t per year.

One of the first industrial-scale oriented investigations of the conversion of CO_2 to methanol was presented in the beginning of the 1990s by Lurgi AG which developed a two-stage process (Fig. 9).[191,194] The make-up gas is pre-converted (10–30% of carbon oxides) in an adiabatic fixed bed reactor in a once-through operation before entering a synthesis loop incorporating a steam raising quasi-isothermal reactor. In the adiabatic reactor the temperature is only slightly increased by maximum 35 °C as the exothermic methanol formation and the endothermic water gas shift reaction proceed simultaneously. Water and methanol formed in the adiabatic reactor are separated, promoting the methanol formation in the second reactor. The overall carbon conversion to methanol is 58–61% (based on data of ref. 194). Results of a single-stage CO_2 to methanol pilot plant presented by Lurgi (now part of Air Liquide) in 2011[126] showed a total CO_2 conversion of 94–97% (per pass conversion of 30–45%).



	Reactor 1		Reactor 2	
Stream	1	2	3	4
CO_2 (mol%)	23.9	18.5	14.0	11.0
CO (mol%)	0.3	3.6	3.0	2.5
H_2 (mol%)	73.9	64.5	69.1	62.0
CH_4 (mol%)	1.2	1.3	8.4	9.3
N_2 (mol%)	0.7	0.8	5.1	5.6
CH_3OH (mol%)	0.0	4.0	0.3	5.2
H_2O (mol%)	0.0	7.3	0.1	4.1
T (°C)	250	286	240	n.a.
p (bar)	80		78	
GHSV (h^{-1})	11,000		12,000	
Catalyst (kg)	200		800	

Fig. 9 Simplified scheme and process parameters of the Lurgi two-stage process for methanol synthesis from CO_2 (own illustration based on König and Göhna[194]).

Specht *et al.* reported the successful conversion of atmospheric CO_2 (sequestrated by caustic scrubber and electro dialysis) and hydrogen to methanol in a bench-scale test plant in 1996.[195] Per pass conversion of 23%[189] is stated for the methanol converter and a total conversion of 98% is estimated by recycling unreacted syngas.[123]

The CAMERE process (carbon dioxide hydrogenation to form methanol *via* a reverse water–gas shift reaction) is another two-stage concept that has been developed at the Korean Institute of Science and Technology (KIST). In contrast to the Lurgi process, the first reactor focuses on the endothermic reverse water–gas shift reaction only. It is electrically heated and operated at higher temperatures of 600–700 °C with the goal of 60% conversion of CO_2 to CO.[190,196] Due to the increased CO content and removal of by-product water prior to the methanol synthesis loop, the per pass methanol yield is increased and recycle gas ratio is reduced. In minipilot experiments a methanol yield of 53% (own calculations based on ref. 190) was achieved. Although a significant increase in methanol production is claimed, the STY of 0.6 kg l_{cat}^{-1} h^{-1} 190 (excluding the catalyst volume of the RWGS reactor) is in line with the results from direct hydrogenation pilot plants.[126,189,197,198] A 75 kg per d pilot plant has been constructed in collaboration with POSCO (Korean Pohang Iron and Steel Company) and KEPRI (Korea Electric Power Institute) in combination with a pilot plant for CO_2 separation from a power plant. The pilot plant reached a methanol yield of 70% at a CO_2 conversion in the RWGS reactor of 35%.[196]

A research group of RITE (Research Institute of Innovative Technology for the Earth) and NITE (National Institute for Resources and Environment) conducted short and long-term experiments of in-house catalysts with a 50 kg d^{-1} oil-cooled test plant.[112,116,184,192] Based on the Toyo MRF-Z converter, a double-train 8000 t per d methanol synthesis plant was designed showing that the size of the reactor is almost the same as for natural gas based plants.[116]

Mitsui Chemicals, an industrial partner of the CO_2 utilisation project led by RITE, continued with the catalyst development and constructed a 100 t per a pilot plant in 2009 based on the established fixed bed synthesis loop scheme.[209,210] It converts by-product hydrogen and CO_2 separated from the exhaust gas both originating from naphtha cracking for an adjacent ethylene oxide unit.[200] For future applications hydrogen production from photo-catalysis[211] and biomass[212] are investigated. Low-cost supply of renewable hydrogen is currently identified to be the major hurdle for commercialisation.[212]

Northern Arizona University has developed a laboratory-scale mobile methanol synthesis test rig for the conversion of decentralised CO_2-sources to methanol. It includes a 30 bar pressurised electrolyser (1.4 kW) for on-site production of hydrogen. The overall efficiency from electricity to methanol (LHV) of first experiments is very poor with only 16.5%. The methanol yield is stated to be 40% although a commercial catalyst has been used.[202]

Silicon Fire AG (Switzerland) has developed a modular container concept, which they claim is ready for series production. The idea is that the container, which includes an electrolyser and methanol synthesis unit, can be placed close to renewable energy sources *e.g.* wind parks, or close to CO_2-sources. More details about the design have not been reported.[203,204]

Carbon Recycling International (CRI) commissioned a carbon dioxide to methanol plant with a capacity of 1300 t methanol per year in 2012 which was expanded to 4000 t per a in 2014.[193] The plant is based on state of the art technology using Cu/ZnO-catalysts operated at 250 °C and 100 atm.[206] The commercial plant is located in Svartsengi, Iceland and recycles 5600 t per a of CO_2 released by a nearby geothermal power plant. The H_2S containing geothermal flue gas has to be desulphurised first.[213] Hydrogen is produced by an alkaline electrolysis unit with a production capacity of 1200 N m^3 h^{-1} (6 MW_{el}).[193]

In 2015 CRI and MHPSE (Mitsubishi Hitachi Power Systems Europe) announced a strategic partnership to offer industrial solutions for carbon capture and power-to-fuel production.[207] Within the EU Horizon 2020 project MefCO2, a demonstration plant was set up and opened in 2019 (1 MW_{el}, 500 t per a CO_2, 400 t per a MeOH).[207,208,214] CO_2 is sequestrated from a coal power plant in Niederaussem (Germany) and hydrogen is produced by electrolysis (supplied by Hydrogenics).

Blue Fuel Energy (Canada)[215] plans to integrate hydrogen and oxygen from electrolysis in a large-scale methanol-to-gasoline plant (2.5 million litres of gasoline per day) based on natural gas reforming. Blue Fuel Energy and Siemens entered a memorandum of understatement for the supply of a 20 MW PEM electrolysis system in 2014. A final investment decision was announced for the end of 2016 but no further information have been reported until now.

One main motivation for power-to-methanol (PtMeOH) represents the CO_2 abatement by CO_2 recycling. Based on the stoichiometry of methanol formation from CO_2 (2), 1.37 t_{CO_2}/t_{MeOH} are utilized. Assuming an overall carbon conversion of 96%, this results in a CO_2 demand of 1.43 t_{CO_2}/t_{MeOH}. The CO_2 emissions of the conventional production process *via* natural gas reforming or coal gasification correspond to 0.52 and 2.83 t_{CO_2}/t_{MeOH} respectively.[216] According to Rivera-Tinoco *et al.* the weighted average emissions of conventional synthesis plants in Europe is 0.77 t_{CO_2}/t_{MeOH}.[217] A lifecycle analysis for Blue Fuel Energy indicates an 84.3% reduction of emissions of PtMeOH compared to gasoline and 81% compared to natural gas based methanol.[218] The overall emissions were determined to be 14.3 g_{CO_2eq} MJ^{-1} (0.28 t_{CO_2}/t_{MeOH}) for PtMeOH using wind energy (including CO_2 separation) and combustion of the methanol as a fuel. The methanol produced in the commercial plant of CRI received certification for a 90% reduction of CO_2 emissions compared to fossil fuels according to the EU Renewable Energy Directive.[193] The specific emissions were determined to be 8.5 g_{CO_2eq} MJ^{-1} (0.17 t_{CO_2}/t_{MeOH}) based on electricity from the Icelandic grid and including transportation by ship from Iceland to Rotterdam.

2.2.2. Flexibility and dynamic operation. PtX is seen as a flexibility option to balance volatile feed-in of wind and PV in a future renewable energy system. Therefore, at least the electrolysis has to be operated dynamically. Ultimately, in order to reduce or avoid large hydrogen intermediate storage, the methanol synthesis needs to be operated in a flexible manner as well.

Steam raising converters are claimed to be very flexible by Lurgi AG with a minimal part-load of 10–15% of the design capacity and a load change rate from zero to full load within a few minutes.[133] Fast load changes are possible due to the large thermal capacity of the surrounding water ensuring a constant behaviour of the reactor. Easy start-up is realised by steam injection initiating the natural water circulation of the cooling cycle and resulting in a uniform heating of the reactor. Dynamic simulation of the two-stage process for methanol synthesis from CO_2 (Fig. 9) by Air Liquide within the project VIteSSe^2 implied a transient production rate step between 20–100% within 6–7 minutes.[219]

The LPMeOH converter was demonstrated to be suited for load-following operation with variable feed flows (5% of design flow per minute) and compositions as well as on/off operation.[164]

For CO_2 operation, Göhna and König observed an activation phase after start-up and concluded that there is a reversible change in the catalyst caused by the reagents or the products.[191] It takes about 60 h (at 95 bar) to 80 h (at 80 bar) starting from about 90% of the full activity until steady state is reached. This was also reported by Pontzen *et al.* using a catalyst from Südchemie.[126] The slow increase of conversion after restart has to be investigated in more detail and has to be taken into account for flexible operation.

2.2.3. Techno-economic analysis. PtMeOH plants have been analysed in several publications. Anicic *et al.* compared direct CO_2 hydrogenation with the CAMERE process.[220] Kiss *et al.* presented an innovative concept of methanol production from CO_2 and wet hydrogen (saturated with water) which is available as a by-product from chlor-alkali electrolysis.[221] Water is removed from the wet-hydrogen while taking CO_x from the methanol–water mixture by stripping the condensed methanol–water mixture from the flash separation. Several energetic analyses of PtMeOH reference plants were published,[24,222–224] as well as techno-economic analyses[189,217,225–228] partly taking CO_2 separation into account. Techno-economic analyses of methanol synthesis for CCU are performed by Tremel *et al.*,[78] Pérez-Fortes *et al.*[229] and Asif *et al.*[230] PtMeOH concepts integrated in biomass gasification were analysed by Clausen *et al.* and Specht *et al.* for alkaline electrolysis[225,231] and by Lebæk *et al.* for SOEC.[232] In addition life-cycle analysis of Power-to-Methanol reference plants have been conducted.[218,233] CRI and Mitsubishi Hitachi Power System analysed the integration of PtMeOH in a steel mill.[234,235]

Taking the thermal energy penalty of CO_2 separation into account, the efficiency from PtMeOH with low temperature electrolysis is in the range of 38% from atmospheric CO_2[225] and 46–48% from flue gas derived CO_2.[15,218,222,225] Mignard *et al.*[228] report higher efficiencies of 52–58% due to a high electrolysis efficiency of 72% and temporally part-load operation (with 81% at 20% load). The assumed or simulated performance of the electrolysis has a major impact on overall efficiency as electrolysis represents by far the major consumer (97% of electric energy demand).[222] High temperature electrolysis promises higher efficiencies. Rivera-Tinoco *et al.*[217] showed an improvement of overall efficiency of 9.5 percentage-points for SOEC (54.8%) compared to PEM electrolysis. However, the efficiency of the PEM-electrolysis based process not taking CO_2 separation into account is low with 45.3% compared to 49–52% determined by others[24,216,222] and 61% claimed by CRI and Mitsubishi Hitachi Power Systems.[193,235] Integrated concepts offer an efficiency potential due to the integration of waste heat streams (*e.g.* for methanol distillation), utilisation of by-product oxygen from electrolysis (*e.g.* in a blast furnace[234] or in a gasifier[225,231]) or avoidance of CO_2 separation in the case of biomass or coal gasification.[225]

The major economic assumptions and determined production costs of some studies are compared in Table 12. The given

Table 12 Comparison of economic assumptions and methanol production prices

Description	Capacity (t_{MEOH} d^{-1})	CAPEX[a] (10^3 € per t_{MeOH} per d)	Capacity factor	Electricity price (€ per MW per h)	CO_2 cost (€ per t)	Methanol price (€ per t)	Ref.
LTEL + BM	900	210–290[a1]	91	40	15	252–316	231
LTEL + CO_2	900	380[a2]	91	40	15	555	
SOEC + BM	1053	325[a1]	91	61/82[b]	—[c]	325–375[d]	232
LTEL + CO_2 (flue gas)	2400	393[a2]	88	n.a.[e]	35	294	238
LTEL + CO_2 (flue gas)	178	410[a2]	91	25	—[c]	515	228
CCU-plant	1300	181[a3]	91	95.1	0	724	229
SOEC (20 MW) + CO_2	50[f]	2350[a2]	91	20–50	3–10	700[g]–5460	217
PEM (24 MW) + CO_2	50[f]	390[a2]	92	20–50	3–10	404[g]–890	
LTEL + CO_2	140	330[a2]	68	93	50	980	78
LTEL + CO_2 (flue gas)	n.a.	n.a.	95	26[h]	n.a.	525[h]	225, 189 and 239
LTEL + CO_2 (atm.)	170	355[a2,h]	95	26[h]	—[c]	715[h]	225, 189 and 239
LTEL + BM	53	410[a1,h]	95	26[h]	—[c]	385[h]	225, 189 and 239
LTEL + CO_2	1485	n.a.	92	10–50	n.a.	400–820	24
LTEL + CO_2	1486	n.a.	46	10–50	n.a.	700–1100	
LTEL (140 MW) + CO_2	n.a.	n.a.	85	50	—[c]	913	226

LTEL, low temperature electrolysis, BM Biomass gasification. Unit conversion: 19.9 MJ kg_{MeOH}^{-1}, 0.792 kg l_{MeOH}^{-1}. [a] Based on given total investment costs and capacity, (a1) including CO_2 separation; (a2) CO_2 provided externally; (a3) CO_2 and H_2 provided externally. [b] Average spot price of used electricity price duration curve. [c] CO_2 supply included in CAPEX and OPEX. [d] Based on given prices in $$ per barrel with 1.3 €/$$ in 2011. [e] Supplied from hydropower plant. [f] Estimated based on given efficiency. [g] Including a 50% cost reduction of the electrolyser (2940 € per kW for SOEC, 300 € per kW for PEM). [h] 1956 DM per €.

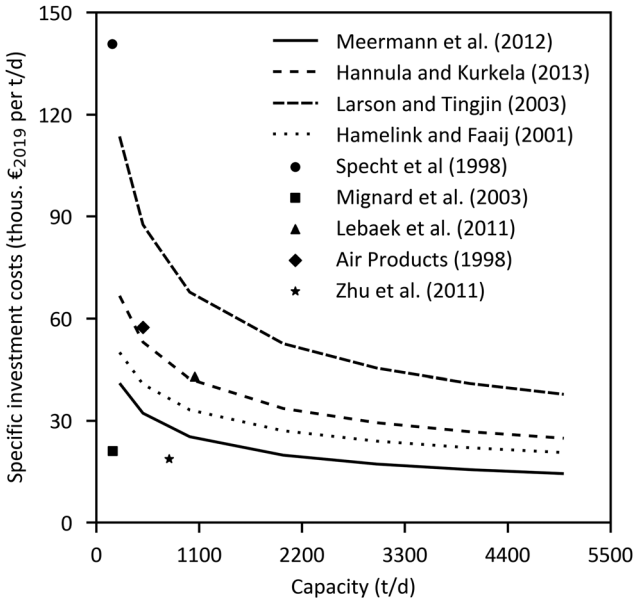


Fig. 10 Specific investment cost curves of methanol synthesis loop including purification normalized to the year 2019.[228,232,239–245]

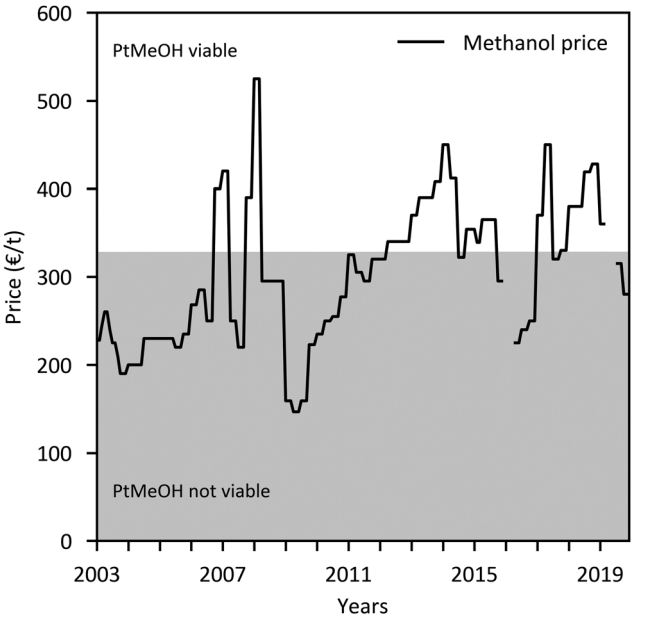


Fig. 11 Historical European methanol prices[236] with price area where PtMeOH might be economically viable (based on the featured techno economic analyses).

specific CAPEX of PtMeOH plants excluding CO_2 separation unit is in the range of 330–410 € per kg_{MeOH} production per d (except for the SOEC case[217]). Specific investment costs for methanol synthesis and purification from literature (including NG and coal studies) are compared in Fig. 10. One major assumption regarding the methanol production price represents the capacity utilisation. Based on a capacity factor of more than 85% (>7450 h per a) in most cases, a specific price in the range of 300–900 € per t_{MeOH} is determined. It is mainly affected by the electricity price. The market price of methanol was highly variable in the past ranging from 150 € per t to 525 € per t in the last years (Fig. 11), with prices of around 260 € per t at the beginning of 2020.[236] Regional benchmark methanol production costs were identified by Boulamanti and Moya[237] to vary between 51 € per t_{MeOH} in Saudi Arabia and 408 € per t_{MeOH} in Europe in 2013 mainly dependent on the regional feedstock costs.

2.3. Discussion

Methanol production from CO_2 and hydrogen was already conducted at the beginning of industrial production in the 1920's. More intensive research began in the 1990's. However, information from pilot plants is still scarce. Despite of some laboratory and pilot plant tests, long-term pilot plant studies on catalyst deactivation and

on the optimal operating conditions are still missing. Further, new catalysts need to be developed to increase the methanol yield from CO_2 hydrogenation and to reduce problems caused by increased water production. In addition, efforts to develop reactor concepts that improve the equilibrium yield of the process should be intensified. Due to less by-product formation, cost and energy savings for the product upgrade need to be considered as well.

Currently, commercial converter designs are largely influenced by the necessity for efficient heat removal. The lower heat released by the methanol synthesis from CO_2 can enable more simple reactor designs and therefore also reduce the complexity of the process design in general.

For state of the art methanol processes a trend towards larger-scale plants to benefit from economy of scale effects exists. Though, using PtMeOH as CCU technology or an operation close to renewable energy sources will require profitable small-scale plants. Consequently, additional efforts are necessary to adapt state of the art reactor and process designs for small-scale applications and develop new concepts better suited for a decentralized and flexible operation.

The identification of new business cases with concepts integrated in other processes and low cost supply of hydrogen play a crucial role in the commercialisation of CO_2 utilisation-based methanol plants. In addition, the feasibility of dynamic operation has to be investigated, to make use of electricity price variations and to act as an energy storage in renewable energy systems.

3. DME synthesis

3.1. State of the art DME synthesis

3.1.1. Theory. DME used to be a by-product of the high-pressure methanol synthesis.[246] The replacement of the high-pressure process by the low-pressure process and the rising demand of DME led to special DME synthesis plants. Currently, DME is produced exclusively by a two-step (or indirect) process of methanol production from syngas followed by dehydration of methanol according to the reaction:

$$2CH_3OH \rightleftharpoons CH_3OCH_3 + H_2O \quad \Delta H = -23.4 \text{ kJ mol}^{-1} \qquad (5)$$

However, DME can also be produced directly from syngas. In this case, methanol synthesis reaction (4), water gas shift reaction (3) and methanol dehydration reaction (5) take place simultaneously and form a synergistic system.[88,247–249] On the one hand, the dehydration or consumption of methanol relieves the equilibrium constraints of methanol synthesis. On the other hand, water formed in reaction (5) is consumed to some extent by the water gas shift reaction and hydrogen is produced which increases the rate of methanol synthesis. Additionally, the consumption of water protects the catalyst from degradation due to water accumulation.[250] The enhanced syngas conversion compared to the methanol synthesis reaction is shown in Fig. 12 for the two overall reaction routes from syngas:

$$3CO + 3H_2 \rightleftharpoons CH_3OCH_3 + CO_2 \quad \Delta H = -246.0 \text{ kJ mol}^{-1} \qquad (6)$$

$$2CO + 4H_2 \rightleftharpoons CH_3OCH_3 + H_2O \quad \Delta H = -205.0 \text{ kJ mol}^{-1} \qquad (7)$$

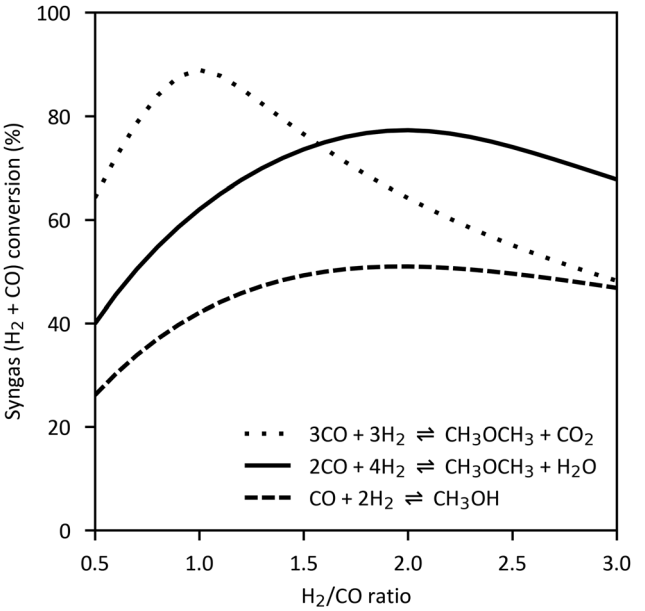


Fig. 12 Syngas conversion for methanol synthesis and DME synthesis reactions as function of H_2/CO ratio (260 °C, 50 bar, own illustration of ref. 30 based on Aspen Plus Simulation).

The direct process of Haldor Topsøe follows reaction (7) while the processes of JFE, Air Products and KOGAS follow reaction (6) resulting in a H_2O or CO_2-rich product respectively. This poses different requirements for the product upgrade discussed later. Based on the reaction stoichiometry, the H_2/CO ratio should be adjusted to 1.0 for reaction (6) and to 2.0 for reaction (7). As the reactions (6) and (7) are exothermic and reduce the number of moles, direct DME synthesis is favoured by low temperature and high pressure. Operation conditions of direct DME synthesis plants have a temperature range of 200–300 °C and a pressure range of 30–70 bar.

In case of CO_2-based DME synthesis, the optimal stoichiometry for direct hydrogenation is given by a H_2/CO_2 ratio of 3 to 1 following on from the stoichiometry of the combination of methanol synthesis reaction (2) and DME synthesis reaction (5). Equilibrium limited conversion of CO_2 and selectivity of DME decreases at a higher temperature and a lower pressure while CO selectivity increases due to the RWGS reaction as shown in Fig. 13.

3.1.2. Catalysts. For the dehydration of methanol according to reaction (5), solid acid catalysts are used, predominantly γ-alumina and zeolites such as HZSM-5. Other catalysts applied or investigated are for example, silica/alumina, aluminum phosphate and ion exchange resins.[30,32,251] For a more detailed overview of catalysts, please see the extensive review of DME catalysts and preparation methods given by Azizi *et al.*[32] and Sun *et al.*[252]

The direct conversion of syngas to DME is realised by using a hybrid or supported bifunctional catalyst. Hybrid catalysts are prepared by physical mixing of powder or granules of conventional methanol synthesis (*e.g.* $CuO/ZnO/Al_2O_3$) and dehydration catalysts (*e.g.* γ-alumina).[30,32] Supported bifunctional catalysts commonly comprise Cu/ZnO as a metal side and dehydration catalyst as a support.[252] The methanol synthesis

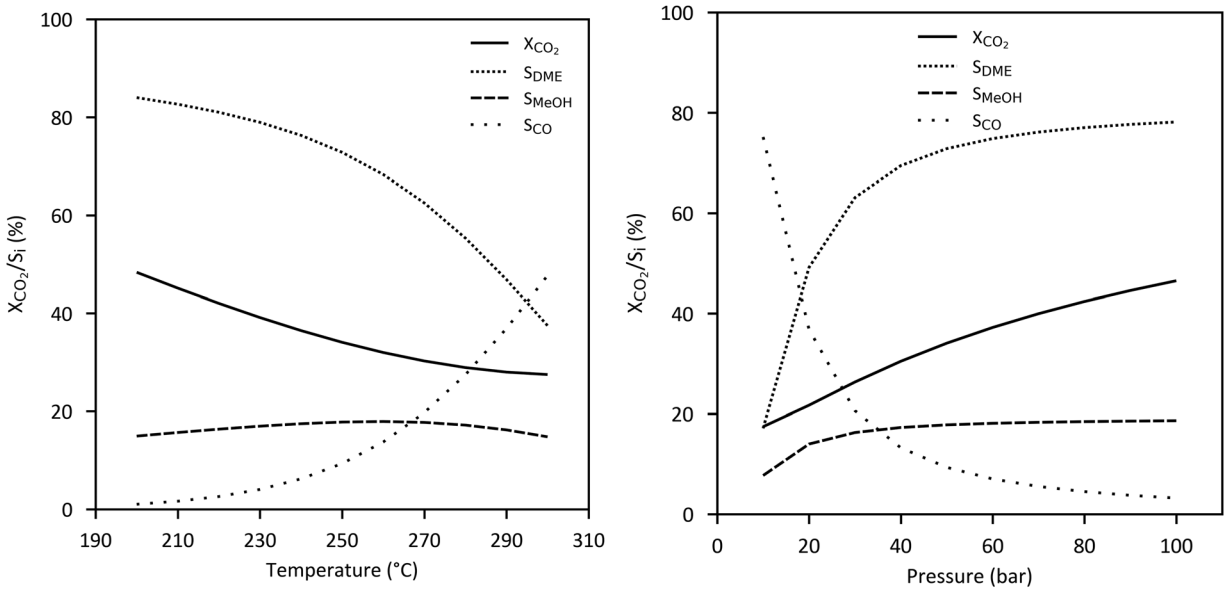


Fig. 13 Equilibrium conversion of CO_2 (= $[CO_{2,in} - CO_{2,out}]/CO_{2,in}$) and selectivity of DME, methanol and CO ($S_i = i/[CO_{2,in} - CO_{2,out}]$, with i = 2 × DME/MeOH/CO, based on amount of C-atoms) dependent on (a) temperature and (b) pressure ($H_2 : CO_2$ = 3 : 1, based on Aspen Plus Simulation).

reaction (2) and the WGS reaction (3) are catalysed by the methanol synthesis catalyst. The dehydration reaction (5) is catalysed by the acidic component. As a result, the product ratio of methanol and DME can be adjusted by the ratio of dehydration and methanol synthesis catalyst, which represents an important design parameter also with regard to the consumption and production of water.[88] Catalyst mixtures applied for the production of DME as the main product contain 33–80 wt% DME catalyst (mainly ranging from 33–50 wt%).[88,251,253–256]

According to MGC, the expected life time of the dehydration catalyst in a commercial two-step process is 4–6 years or more, comparable to the methanol synthesis catalyst.[30] ENN reports a life time of their dehydration catalyst of more than 24 months.[257] Haldor Topsøe operated a 50 kg per d direct DME synthesis bench plant with a hybrid catalyst charge for more than 14 000 h.[130,249] The results indicate higher stability compared to their commercial methanol synthesis catalyst MK-101, which was in operation in a large methanol plant for more than 5 years. For the JFE direct DME synthesis process it is expected that the catalyst life is at least one year for a commercial scale plant, maintaining DME production by gradually increasing the operation temperature.[258]

Deactivation of dehydration and hybrid catalysts is caused by sintering active copper sites, coke deposition, poisoning by contaminants in the syngas or adsorption of water (also formed from CO_2) and thus blockage of acidic sites.[259] The requirements of the syngas purity to avoid deactivation by contaminants is specified in Table 13. While Wang *et al.* reported that the deactivation of hybrid catalysts is caused mainly by the deactivation of the Cu-based methanol catalyst,[260] results from Air Products indicate an almost similar deactivation rate of methanol catalyst and gamma alumina (γ-Al_2O_3).[261] γ-Al_2O_3 and Cu-based methanol synthesis catalysts show a higher degradation in slurry reactors than fixed bed reactors which represent the main reactor types of direct DME synthesis.[32] According to Fujimoto *et al.*[30] the rate of water transfer from alumina to the methanol catalyst is much lower in the slurry phase reaction than in a fixed bed system. However, the problem of increased deactivation could be solved by adding an active component, which promotes the shift reaction and thus efficiently removes water as hydrogen from the alumina surface. Moreover, in the beginning of the direct DME synthesis development of Air Products (LPDME), rapid degradation of γ-Al_2O_3 and methanol synthesis catalyst in a physical mixture was observed due to cross metal contamination.[262] A comparison of degradation over time of a catalyst from Haldor Topsøe and JFW for direct synthesis is shown in Fig. 14. As seen, in the first 1000 h, deactivation is in the range of 10–25% of the initial value.

Table 13 Syngas specification for DME synthesis[30]

Component	Purity requirement
NH_3, HCN	<1.0 ppm (vol)
H_2S, COS, CS_2, Fe carbonyl, Ni carbonyl, Cl, F	<0.1 ppm (vol)

3.1.3. Commercial converter designs and processes. Currently, there are several licensors that offer large-scale DME synthesis plants based on the direct as well as two-step process. All commercial large-scale DME synthesis plants so far are based on the two-step process despite the thermodynamic advantages of the direct process.

The two-step process benefits from the simple extension of the well-established methanol process available on a very large scale associated with low additional investment costs. Moreover the two-step process offers advantages in product upgrade discussed later and the possibility to change the product ratio of methanol and DME according to the market demand.[30,263]

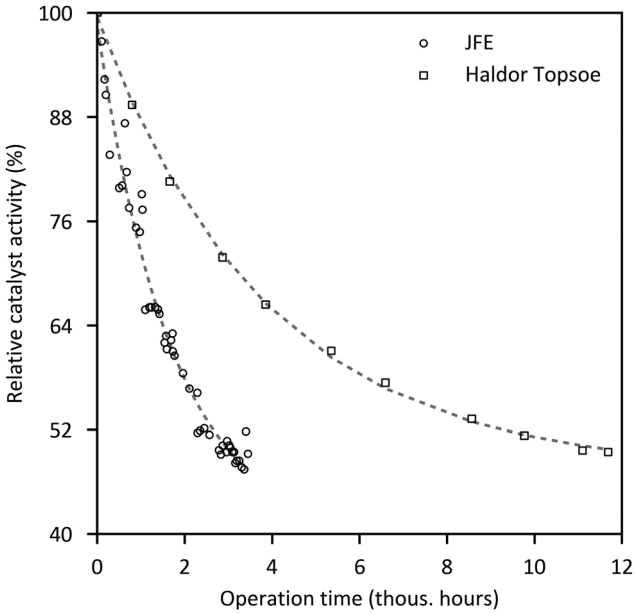


Fig. 14 Deactivation of direct DME synthesis catalysts of Haldor Topsøe[130] and JFE (260 °C, 50 bar).[258]

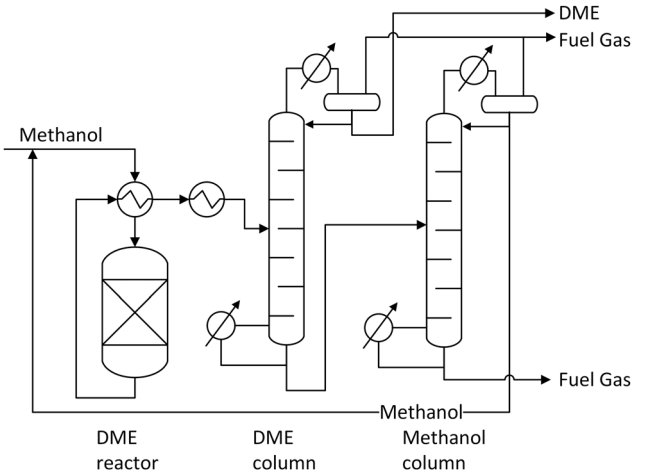


Fig. 15 Typical flowsheet of the DME section of a two-step synthesis process.

Lurgi claims that the investment costs of the two-step process including product upgrades are 10–20% lower on a 3000–5000 t per d scale than the direct synthesis process.[264]

Most methanol technology licensors like Haldor Topsøe, Lurgi, Toyo Engineering Corporation (TEC), Mitsubishi Gas Chemical or Johnson Matthey Technology also offer two step DME processes up to the scale of their methanol technology. For example, TEC states that a single train JumboDME plant can be realised based on their JumboMethanol scheme with a DME capacity of 3500 t per d and a radial reactor similar to the MRF-Z converter with a capacity up to 7000 t per d is feasible.

Beside the established companies there are some emerging Chinese companies like China Energy,[265] Tianyi[266] or ENN.[267] A typical flowsheet of a two-step DME synthesis is shown in Fig. 15. Stabilised methanol (after the first distillation column), further upgraded methanol (Grade A or AA) or even raw methanol directly after the methanol converter[268,269] are used as feedstock. The methanol is vaporized and fed to the dehydration reactor where it is partly converted to DME, water and very small amounts of light ends (CH_4, CO, CO_2, H_2). Methanol consumption is about 1.4 $t_{Methanol}/t_{DME}$, which is close to the theoretical minimum of 1.39 $t_{Methanol}/t_{DME}$. Unconverted methanol is separated from the DME and recycled to the DME reactor. Per pass conversion is in the range of 70–85% at typical operating conditions of 220–400 °C and a low pressure of 1–30 bar (see Table 14). Apart from China Energy all technologies are based on methanol dehydration in a fixed bed reactor.

Direct DME synthesis processes have been developed and demonstrated on a pilot plant scale by JFE, KOGAS, Haldor Topsøe and Air Products and Chemicals Inc. These processes are compared in Table 15 and will be discussed in the following in more detail.

Haldor Topsøe have tested their self-developed dual function catalyst (*e.g.* aluminosilicate catalyst)[282] in a 50 kg per d cooled reactor pilot plant as presented in 1995 by Hansen *et al.* showing a comparable durability to their methanol catalyst MK-101.[249] Laboratory experiments at 37 bar and 220–330 °C were reported by Hansen and Joensen in 1991.[171] These experiments indicated a very low amount of by-products even with hot spot temperatures above 300 °C. The Haldor Topsøe process was based on reaction (7) resulting in a product with a high water content. For large-scale applications, three adiabatic reactors with interstage cooling are proposed by Dybkjaer and Hansen.[130] This enables single train capacities of up to 10 000 t per d methanol equivalent. However, Haldor Topsøe's commercial projects are all based on their two-step process presented before.

KOGAS (Korea Gas Corporation) began the development of DME synthesis from natural gas based on a shell and tube steam raising converter design in 2000.[283] The catalyst is filled in tubes surrounded by boiling water on the shell-side. In addition, KOGAS examined a mixture of methanol and dehydration catalyst pellets and developed a proprietary hybrid catalyst. A higher activity and DME selectivity of the hybrid catalyst was shown in laboratory experiments as well as reactor simulations.[251,256] This proprietary hybrid catalyst consists of a mixture of fine powdered methanol catalyst ($CuO/ZnO/Al_2O_3$ with added promoters Mg, Zr, Ga, Ca) and dehydration catalyst (γ-alumina with aluminum phosphate).[251,255] In 2003, a 50 kg per d pilot plant was set-up and a 10 t per d demonstration plant was brought into operation in combination with a natural gas tri-reformer in 2008. Based on flowsheet modelling[284] and numerical simulations of the reactor,[253] a conceptual design of 3000 t per d DME process was completed by 2009.[283,285] However, problems with degradation during the demo plant tests were reported due to hot spots.[30] As a result, for the 3000 t per d concept the CO_2 content was planned to be 25–30%, resulting in a lower conversion and a reduced increase in temperature.

Table 14 Comparison of commercial two-step processes

Company	MGC	Lurgi (air liquide)	Toyo (TEC)	Haldor Topsøe	ENN
Methanol	Grade AA	Stabilised methanol[a]	Stabilized methanol	Crude, fuel grade, grade A	>50% MeOH
Reactor design	Adiabatic fixed bed	(Adiabatic or cooled) fixed bed	Adiabatic fixed bed (radial)	Adiabatic fixed bed	(Adiabatic or cooled) fixed bed
p (bar)	10–25	1–30	10–20	10–15	6–12
T (°C)	250–400	250–360	220–250 (inlet), 300–350 (outlet)	290–400	260–360 (1st stage), 180–240 (2nd stage).
Per pass conversion	70–80%	70–85%	70–85%	80%	>80%[c]
Selectivity	98.60%	n.a.	99.9%	n.a.	>99%
Catalyst	γ-Alumina	n.a.	γ-Alumina	Activated alumina[b]	Alumina-silica
Plants	Niigata (80 ktpy), Trinidad and Tobaga (20 ktpy)	—	4 plants in China (total 470 ktpy)	BioDME Pitea (5 tpdy), Iran (800 ktpy)	Zhangjjagang (200 ktpy), Bengbu (20 ktpy)
Ref.	30, 179 and 270	126 and 271–273	263 and 274–277	167–169, 249, 278 and 279	257, 267, 280 and 281

[a] <50 wt ppm carbonyl compounds, about 20% water. [b] Trademark DK-500 and DME-99 Eco. [c] Two-stage DME conversion process. Per pass conversion is 70–80% in the first stage and 85–92% in the second stage.

Table 15 Overview of direct DME synthesis technologies

	Haldor Topsøe	KOGAS	JFE	Air products
H_2/CO	2	0.8–1.5	1	0.5–0.7
Reactor design	Adiabatic fixed bed	Tubular SRC	Slurry	Slurry
p (bar)	37–42	30–60	50 (30–70)	52
T (°C)	240–330	260 (200–300)	260 (240–280)	250
GHSV	n.a.	2000–10 000 h^{-1}	4000 sl per kg_{cat} per h (3000–8000)[e]	6000 sl per kg_{cat} per h (5700/9000)
Catalyst MeOH/DME	Cu-based/aluminosilicate	$CuO/ZnO/Al_2O_3$ $AlPO_4$/γ-alumina	$CuO/ZnO/Al_2O_3$ γ-alumina	$CuO/ZnO/Al_2O_3$[g] γ-alumina
Catalyst ratio (MeOH : DME)	n.a.	0.8–4 : 1	2 : 1 (0.75–20 : 1)[f]	95:5 (80 : 20)
Product (DME purity %)	MeOH/DME mixture	99.6	>99.5% (99.9)	MeOH/DME mixture
Recycle ratio	4–5[c]	n.a.	1.6–1.8	n.a.
CO conversion (per pass)	40–70[c]	67[d]	50–60	19–24 (21–31)[h]
Overall CO conversion (%)	94–97[c]	n.a.	94–96	n.a.
H_2O selectivity[a] (%)	18–28[c]	n.a.	1.3	<1[h]
DME selectivity[b] (%)	74–76[c]	60–70	91	61–67 (40–80)[h]
Plants	1993: 50 kg per d Bench	2003: 50 kg per d Bench 2008: 10 t per d Demo	1995: 50 kg per d Bench 1997: 5 t per d Pilot 2003: 100 t per d Demo	1991: 4–6 t per d Pilot 1999: 8–9 t per d Demo
Ref.	130, 171 and 301	251, 255, 256, 283, 284 and 300	30, 250, 254, 258 and 295–299	261, 288 and 294

[a] Product C-mol ratio: 2 × DME/(2 × DME + MeOH). [b] Product H-mol ratio: 2 × H_2O/(2 × H_2O + 6 × DME + 4 × MeOH). [c] Calculation based on patent WO96/23755.[301] [d] 65–82% dependent on CO_2 in feed gas.[31] [e] About 6 g_{cat} h mol^{-1} for 100 t per d and 3000 t per d concept,[31] 4 (3–8 g_{cat} h mol^{-1} for 5 t per d plant).[250] [f] Based on patent US 2006/52647[295] when the methanol catalyst substitutes the WGS catalyst. [g] Commercial catalyst/t BASF S3-86. [h] Own calculations based on Air Products and Chemicals Inc.[261,288]

The Japanese Corporation JFE (formally NKK) developed a direct DME synthesis process based on a slurry reactor in various collaborations.[250,258] It started with the development of a new catalyst in 1989.[250] Afterwards, a 50 kg per d bench unit was brought into operation in 1995, followed by a 5 t per d pilot plant in 1997[286] and a 100 t per d demonstration plant (including autothermal reformation of natural gas, process scheme see Fig. 16) which was operated from 2003 to 2006 with a total production of 19 520 t of DME.[30,258] JFE selected the slurry reactor design due to its excellent heat transfer capabilities, resulting in almost isothermal temperature profiles within the reactor.[250,258] The once-through CO conversion is in the range of 50–60% (at standard conditions of 260 °C, 50 bar and about 6 g_{Cat} h^{-1} mol^{-1}).[30,31] The overall CO conversion reaches up to 96% at a recycling ratio of about 1.8.[250] The cold gas efficiency of the DME section in the 100 t per d demonstrations tests was 81.7%.[258] A 3000 t per d reactor concept was presented with an inner diameter of 7 m and a height of 50 m (slurry 46 m), which is a scale-up of the 10 t per d reactor which had an inner diameter of 2.3 m and a height of 22 m (slurry 15 m).[258,287]

Air Products and Chemicals investigated the co-production of DME within the LPMEOH-Project. The proof of concept was demonstrated in 1991 at a production capacity of 4–6 t per d of methanol and DME.[288] A physical mixture of γ-Al_2O_3 (6.6–19.3 wt%) and a commercial methanol synthesis catalyst (BASF S3-86) were used. The CO conversion was identified to be in the range of 21–31%. The DME selectivity is determined to be 40–80% dependent on the amount of γ-Al_2O_3 and the space velocity. In the following, efforts were spent in the investigation of reaction kinetics and improvement of stability of the catalyst

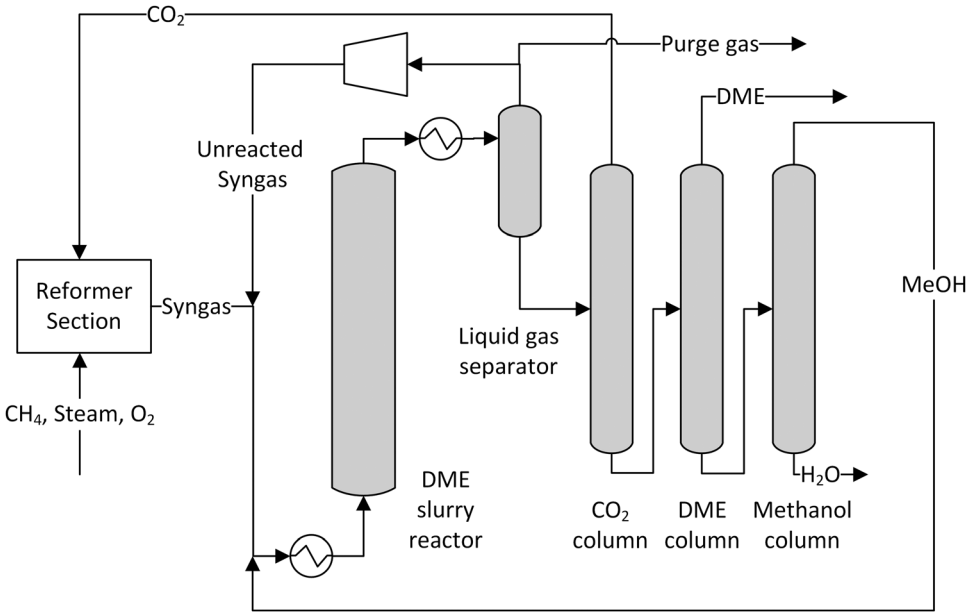


Fig. 16 Schematic process flow of the JFE 100 t per d demonstration plant (own illustration based on Yagi *et al.*[258]).

system.[88,247,248,289] In 1999, the Liquid Phase Dimethyl Ether (LPDME) process was demonstrated in the 10 t per d (4–5 t per d DME) Alternative Fuels Development Unit (AFDU) of the U.S Department of Energy (DOE).[261] A methanol to dehydration catalyst ratio of 95 to 5 by weight was found to be the optimum for catalyst stability. Selectivity was in the range of 61–67%. The deactivation rate for both catalysts was identified to be 0.7% per d. In 2019 BASF, Linde and Lutianhua announced a partnership to build a pilot plant for an energy-efficient one-step DME process.[290] Currently no further details have been reported. A comparison of once-through CO conversion for JFE and LPDME plant concepts is given in Fig. 17.

Although direct DME synthesis is ready for commercialisation and the cold gas efficiency of DME synthesis can be increased by up to 10 percentage-points compared to the two-step process,[30,291] no commercial project has been realised until now. However, due to the high potential, direct DME synthesis still attracts research and commercial interest. For example, Linde AG and BASF SE recently investigated direct DME synthesis technology and catalysts within the project DMEEXCO2 (2012–2015).[291–293]

3.1.4. Product upgrade. In the two-step process, the product upgrade of DME is typically realised by a two-column distillation.[168,271,275] The product stream of the DME reactor is passed to the DME column, where DME and light ends are removed overhead. The light ends are separated from the liquid DME product in a flash and can be used as fuel gas. Unconverted methanol and by-product water are discharged from the bottom of the DME column and passed to the methanol recovery column. The purified methanol is recycled to the DME reactor. A purity higher than 99% can be achieved by this process. To provide DME with a higher purity (>99.9%), Haldor Topsøe proposes to add an olefin stripper prior to the DME column.[302] Generally, the

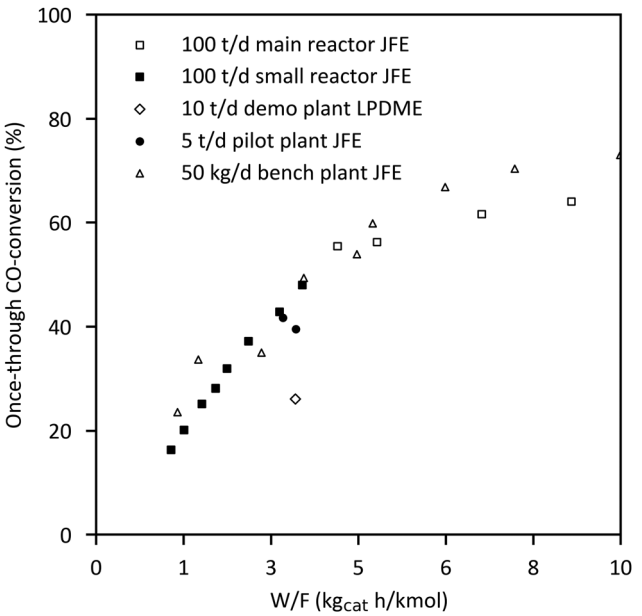


Fig. 17 Dependency of once-through CO conversion on catalyst loading ratio W/F (catalyst weight/reactor inlet gas flow rate) of JFE plants (260 °C, 50 bar)[258] and LPDME 10 t per d demonstration (own calculation based on ref. 261).

product upgrade section features heat integration in various forms as presented by Haldor Topsøe,[167] Toyo[275] or Lurgi.[126] For example, Lurgi couples the methanol vapouriser and the DME-column in its MegaDME concept, so that each becomes the reboiler or overhead condenser of the other.[126,303]

Direct DME synthesis features a more complex product upgrade compared to the two-step process due to the reactor effluent mixture of DME, methanol, CO_2, water and unconverted syngas. The removal of CO_2 especially from the DME product

stream is very difficult. Several ways have been proposed to solve this problem. Examples are the separation of CO_2 by membranes[293] or to avoid the formation of large amounts of CO_2 by operation in a H_2-rich regime (*e.g.* $H_2/CO > 5$).[301]

Because of the low boiling point of DME, a cryogenic separation of liquid products and unconverted syngas is typically applied in the recycle loop.[271] To condense most of the DME along with the methanol and water, the syngas is cooled down to temperatures of about −40 °C in the JFE and KOGAS processes.[284,296] CO_2 is dissolved in the product DME. The unconverted syngas is recycled to the DME synthesis reactor. The liquids are depressurised and directed to the CO_2 column followed by the DME column and the methanol recovery column.[258,284,296] In the KOGAS process, CO_2 is separated from the liquid stream in the CO_2 column operated at a pressure of 35 bar and DME is recovered in the DME column at a pressure of 18 bar.[284]

In a US patent of Air Products,[304] the reactor effluent is fed to a high-pressure flash column. The vapour mixture comprises DME, CO_2 and unconverted syngas. A mixture of DME and methanol is used in a scrubber to remove DME and CO_2 from the unconverted syngas, which is recycled to the DME reactor.

Another concept is disclosed in a US patent of Haldor Topsøe[305] removing the CO_2 in a potassium carbonate scrubber from the reactor effluent. In a subsequent distillation column, methanol and water are discharged in the bottoms and DME is obtained by condensation of the top product. The remaining syngas depleted in CO_2 is recycled to the DME reactor.

3.2. Current research on power-to-DME

3.2.1. DME from CO_2. An overview of possible production pathways of DME from CO_2 is given in Fig. 19. DME can be produced *via* the two-step or direct DME synthesis process from CO_2 or from CO by upstream converting of CO_2 in an additional reverse water gas shift reactor (RWGS).

The extent of synergy of the direct synthesis declines with rising CO_2 content due to the high formation of water *via* the RWGS reaction inhibiting methanol production and the dehydration reaction.[30,88,248] Experiments with varying CO_2 content show a significant decrease of CO conversion.[30,259] Direct hydrogenation of CO_2 to DME has been investigated mainly by using a physical mixture of Cu-based methanol catalysts and zeolites in fixed bed reactors[306–320] as zeolites are not sensitive to water. Hybrid catalysts containing γ-alumina were investigated to a lesser extent.[88,313,321,322] An overview of CO_2 conversion and selectivity of several investigated catalysts is shown in Fig. 18. Per pass CO_2 conversion of maximum 35% (at 210 °C, 50 bar) is achieved. The selectivity of DME lies mainly between 40 and 70%. Naik *et al.*[313] tested two hybrid catalysts (Cu-based catalyst with γ-alumina or HZSM-5) in a fixed bed and slurry reactor. The γ-alumina catalyst shows an inferior performance regarding activity, selectivity and durability compared to the zeolite catalyst. Moreover, CO_2 conversion, DME selectivity and catalyst stability is higher for the fixed bed reactor than for the slurry reactor. This can be explained by water accumulation on the catalyst surface due to the additional mass transfer resistance through the hydrocarbon oil in the slurry reactor. More detail can be found in Alvarez *et al.*, which provide an overview of hybrid catalysts investigated in the last 10 years.[323]

Ateka *et al.*[324] investigated alternative promoters for bifunctional catalysts (Cu-based with Zr or Mn promoters on SAPO-18 zeolite). Zr and Mn metallic functions of the catalyst showed a similar behaviour for the methanol synthesis step, however provided higher yields and selectivity for the DME step. In addition, the catalysts showed a lower deactivation for H_2/CO_2-based feed gas compared to conventional syngas. Catizzone *et al.*

	Catalyst	p (bar)	T (°C)	Type
An et al. (2008)	HZSM-5	50	210-270	FB
An et al. (2008)	HZSM-5	20-50	250	FB
Frusteri et al. (2015)	HZSM-5	30	270	FB
Frusteri et al. (2015)	HZSM-5	30	240	FB
Liu et al. (2013)	HZSM-5	50	240	FB
Liu et al. (2013)	HZSM-5	50	260	FB
Naik et al. (2015)	γ-Al_2O_3	50	260	FB
Naik et al. (2015)	HZSM-5	50	260	SL
Naik et al. (2015)	γ-Al_2O_3	50	260	SL
Naik et al. (2015)	HZSM-5	30	240-280	FB
Sun et al. (2004)	HZSM-5	20-35	260	FB
Wang et al. (2009)	HZSM-5	30	200	FB
Zhang et al. (2014)	HZSM-5	30	250	FB

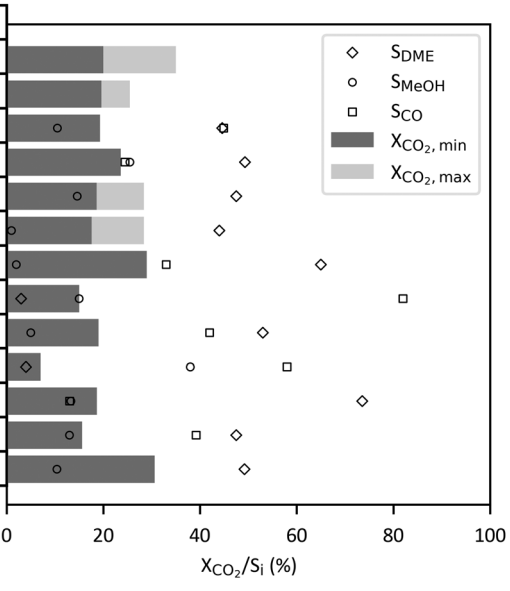


Fig. 18 Overview of CO_2 conversion and selectivity of DME, methanol and CO ($S_i = i/[CO_{2,in} - CO_{2,out}]$, with $i = 2 \times$ DME/MeOH/CO, based on amount of C-atoms) of direct hydrogenation experiments of CO_2 to DME.[306,310,313,314,317–319]

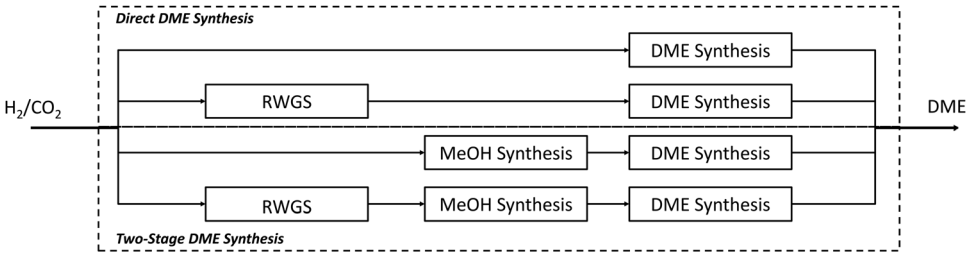


Fig. 19 Possible process pathways from CO_2 to DME.

give a good summary of different bifunctional catalysts that have been investigated recently for direct CO_2 hydrogenation.[325]

A feasibility study of a 500 t per d DME production from CO_2 in Iceland has been commissioned by the Icelandic Ministry of Industry, Mitsubishi Heavy Industries (MHI) and others as part of the long-term vision of a zero percent hydrocarbon fuel emissions society in Iceland.[326] The CO_2 should be captured from the exhaust gas of the ELKEM ferrosilicon plant by MHI's CO_2 recovery process. Hydrogen is supplied by water electrolysis. DME is produced *via* MHI's two-step process from CO_2 and H_2 comprising the superconverter and methanol dehydration technology. The authors conclude that this project is considered to be feasible with dedicated support from the government.

At RWE's Innovation Centre at Niederaussem power station in Germany, a Power-to-X-to-Power research facility was established as part of the multi-partner ALIGN-CCUS project in November 2019. The synthesis of DME investigated at the new pilot plant is based on captured CO_2 from the lignite-fired power plant and electrolysis-based hydrogen. The aimed production rate is up to 50 kg of DME per d in a one-step process.[327] The synthetic energy carrier will then be used and evaluated both for re-electricity generation and use as fuel in vehicles that are difficult to electrify.[328,329]

Matzen and Demirel[330] conducted a life-cycle assessment of the production of renewable methanol and DME (*via* the two-step process) based on Aspen Plus simulations. The use of fermentation-based CO_2 and H_2 from wind-powered electrolysis results in a reduction of greenhouse gas emissions by 86% for methanol and 82% for DME compared to conventional petroleum.

Vibhatavata *et al.* discussed the production of DME from CO_2 emissions of the cement industry and hydrogen produced by nuclear-powered water electrolysis in France. The overall efficiency of the CO_2-to-DME process was 53%, comprising CO_2 capture (MEA wash) from flue gas of the cement industry, reverse water gas shift reactor, electrolysis, methanol and DME synthesis. A similar efficiency in the range of 52–55% from electricity to DME was estimated for a reference case based on electrolysis, RWGS and direct DME synthesis (excluding CO_2 capture effort) by Ohno *et al.*[31,331] For the RWGS autothermal operation without catalyst at a temperature of 1000 °C by adding O_2 from the electrolysis unit is proposed.

Sun *et al.*[332] determined the feasible operating conditions of a pressurised SOEC integrated in a DME synthesis process. The thermodynamic analysis indicates that a high operating temperature of the SOEC is necessary (*e.g.* >900 °C at 80% reactant utilisation and 50 bar) to avoid carbon formation due to the required high pressure and high CO_2/H_2 ratio (≥ 1) of the DME process.

An *et al.*[306] showed that the recycling of unconverted syngas in direct synthesis of DME from CO_2 not only increases overall conversion but also increases space-time yield of DME. This is because the reactor effluent contains CO as a result of the fast water gas shift reaction and the recycling of CO increases methanol and DME yield.

3.2.2. Flexibility and dynamic operation. Research on the flexibility and dynamic operation of DME plants is still scarce. In particular, a comparison of two-stage and direct DME processing in regarding flexibility is missing.

Peng *et al.* investigated the influence of changes of the H_2/CO_2 ratio on DME production for a single step DME synthesis with recycle.[333] They conclude that the DME productivity is very sensitive to changes in the H_2/CO_2 ratio due to an amplifying effect of the recycle.

Gonzáles *et al.* investigated the flexible operation of fixed-bed reactors for CO_2 hydrogenation reactions. DME is identified as a promising route to be operated under flexible conditions due to its high and robust product selectivity. In addition, the limited sensitivity of the DME synthesis towards temperature variations and the high catalyst stability for DME synthesis are seen as beneficial.[334]

Farsi and Jahanmiri modelled and simulated the behaviour of a fixed bed MeOH to DME reactor (second stage of the two-stage DME synthesis). They applied different disturbances and investigated the open loop behaviour of the reactor. Their results show that a 2 bar step change in pressure does not significantly affect the product composition. Feed composition and temperature, however, have a large influence on the product composition.

Hadipour and Sohrabi developed a dynamic model of a direct DME reactor based on experimental data. They conducted experiments at 230–300 °C and 9 bar. The experimental data shows that for all temperatures it takes around 300 min from start up to reach a steady state.[335]

3.2.3. Techno-economic analysis. Extensive techno-economic analyses for Power-to-DME (PtDME) are still scarce.[336] In particular, an economic comparison of the different process routes from CO_2 described in Section 3.2.1 within the PtL context is missing. Several authors have investigated the production of DME from biomass, estimating the manufacturing cost in a range of 302–891 € per t.[337–339] Further, CCU concepts based on coal gasification and natural gas reforming using the direct and

two-stage process route have been analysed,[340–342] indicating cost advantages for the two-stage process route due to the simplified product upgrade.[342] Due to the different boundary conditions and plant concepts (*e.g.* IGCC polygeneration plant, shale gas reforming), a cost comparison is difficult and no economic meaningful recommendation for PtDME processes can be drawn.

While first cost estimates of PtDME differ in their approach and investigated process chain, they agree on DME costs above 1000 € per t, therefore significantly higher than DME from biomass or coal. Michailos *et al.* examine the two-stage synthesis of DME using CO_2 captured from a cement plant and hydrogen from a PEM electrolyser.[343] The simulated plant, designed for a throughput of 740 t of DME per d, can synthesise DME at costs five times the price of conventional diesel (based on the LHV). A Monte Carlo simulation-based sensitivity analysis indicates that the 95% confidence interval for the DME minimal sales price ranges from 1828–2322 € per t.[343]

Schemme *et al.* compared different H_2-based synthetic fuels, indicating DME and MeOH as the fuels with the lowest manufacturing cost. They consider direct DME synthesis from CO_2. Compared to the LHV-based diesel equivalent (DE), they estimate DME manufacturing costs to be 1.49 € per l_{DE} or 1490 € per t, which is in good agreement with Tremel *et al.*,[78] who also considers the direct synthesis of DME and estimates costs of 1390 € per t. The sensitivity analysis shows that after hydrogen and CO_2 costs, the fixed capital investment has the biggest effect on DME costs.[344] In China with its big DME market, prices in 2020 varied between 329 € per t und 276 € per t, leading to a big spread in PtDME production costs.[345,346]

3.3. Discussion

Currently, DME is almost exclusively produced from syngas in a two-stage process with methanol as an intermediate. Consequently, equivalent PtDME processes can benefit from the long-term operational experience. However, an additional RWGS-reactor is necessary, increasing capital expenditures. An alternative two-stage design might include the direct conversion of CO_2-based feed gas to methanol in the first reactor, thus benefitting from experiences from power-to-methanol pilot plants. Both variants offer the advantage that the product ratio of methanol-to-DME can be adjusted based on current market prices. However, this is accompanied by a more complex process design and disadvantages in terms of chemistry.

The one-step process can offer a synergetic effect between the ongoing reactions, however, an increasing CO_2 content in the feed gas reduces these benefits due to higher formation of water. Hence, a shift of CO_2 to CO before the DME reactor might be necessary. Therefore, it is not certain that process complexity decreases with the one-step-process, especially if the more challenging product upgrade is considered. While catalyst development for the direct conversion to DME has already attracted research interest, additional long-term catalyst tests are necessary for a better understanding of catalyst stability and deactivation behaviour. In particular, bifunctional catalysts, which have already shown their high potential regarding selectivity and stability, should be further investigated.

In order to evaluate the different production pathways for PtDME, extensive techno-economic analyses are needed. Different demands determined by the use-case as well as plant location require not only the consideration of classical process factors (such as process pathway, energy consumption or operational expenditures), but in addition the process surroundings (*e.g.* CO_2-source) and market situation (*e.g.* optimal product ratio). Furthermore, a lack of long-term experiments, pilot plant operation and flexible operations are evident to enable a full assessment of the different pathways.

4. Fischer–Tropsch synthesis

4.1. State of the art Fischer–Tropsch-synthesis

4.1.1. Theory. The Fischer–Tropsch (FT) process is a heterogeneously catalysed pathway to convert syngas to liquid hydrocarbons. The products are a variety of simple hydrocarbon chains of different lengths, depending on the process conditions. They can be applied as a low-sulphur diesel substitute for example.

Historically, FT-synthesis was developed in the early 20th century as an alternative fuel production route from coal gasification. After the work on the hydrogenation of CO published by Sabatier in 1902,[347] Fischer and Tropsch worked on the catalytic hydrogenation of CO to various products and published their findings in 1923.[348,349] Soon after its discovery it was recognised that temperature control of the exothermal reaction would become far easier in a liquid phase synthesis. In 1930 the first pilot plant ran in liquid phase modus.[350] The first commercial plant based on a nickel catalyst was constructed in 1933 at Ruhrchemie. In the 1940s, a total capacity of 0.6 million t per a was installed in Germany.[36] All these plants were shut down after WWII due to economic reasons.[36] In the second half of the 20th century, large-scale processes were applied: a multi-tubular fixed bed reactor with gas recycle was developed by Lurgi and Ruhrchemie. Compared to precursor reactors, this one uses increased temperatures and pressures, a stable reaction rate profile over the length and a better heat removal due to higher gas velocities. This process is called Arge process and was the basis for the Sasolburg plant.

Generally, syngas with a H_2/CO ratio of 2–2.2 is processed on a Fe- or Co-catalyst at temperatures of at least 160–200 °C at ambient pressure. Higher pressures and temperatures are advantageous. Reactions (8)–(10) take place during FT-synthesis with n as the resulting carbon chain length (typically ranging between 10 and 20). Reaction (8) is the desired one for the production of alkanes. Water, which is always a by-product, can be converted *via* WGS reaction with CO to CO_2 and H_2 by using Fe-catalysts.[123]

$$n\mathrm{CO} + (2n+1)\mathrm{H_2} \rightleftharpoons \mathrm{C}_n\mathrm{H}_{2n+2} + n\mathrm{H_2O} \quad \Delta H = n(-146.0)\ \mathrm{kJ\ mol^{-1}} \tag{8}$$

$$n\mathrm{CO} + (2n)\mathrm{H_2} \rightleftharpoons \mathrm{C}_n\mathrm{H}_{2n} + n\mathrm{H_2O} \tag{9}$$

$$n\mathrm{CO} + (2n)\mathrm{H_2} \rightleftharpoons \mathrm{C}_n\mathrm{H}_{2n+1}\mathrm{OH} + (n-1)\mathrm{H_2O} \tag{10}$$

The FT-synthesis takes place at temperatures between 150–300 °C. Higher temperatures lead to undesired short

chains and methane, the catalyst may be damaged and carbon deposition can occur. Lower temperatures are limited by reaction velocities and the conversion rate. The pressure can be increased to favour long-chain products and to increase the conversion rate. If the pressure is too high, coke formation can lead to catalyst deactivation and expensive high-pressure equipment is needed. The educt gas composition should have a H_2/CO ratio around 2 for Co-catalysts, Fe-catalysts tolerate higher CO concentrations. Usually, the carbon source on which the FT-process is based on is CO from gasification of coal or other solid feedstock or natural gas. If natural gas is used as a feedstock, steam reforming has to take place to transform the methane molecule to CO and H_2. By using the WGS reaction even CO_2 can be used as a carbon source for the synthesis. Regardless of the carbon source, the desired H_2/CO ratio is set *via* the WGS reaction to the ideal level of two. If no pure H_2 and CO gas mixture is used as feedstock, a separation of sulphur, which is a catalyst poison, must be ensured.

As mentioned above, the products of the FT-process are carbon chains of different length. The chain length distribution of the alkanes is described by the Anderson–Schulz–Flory distribution (11), which is shown in Fig. 20 for different chain length regions.[351,352]

$$W_n = n(1 - \alpha)^2\alpha^{n-1} \quad (11)$$

where W_n denotes the weight fraction of the alkene with the chain length n and α is the chain growth probability, which depends on the catalyst type (Fe or Co), the process temperature and pressure as well as the syngas composition.

Table 16 shows the chain growth probability that is necessary to obtain a certain product range (number of C atoms). Depending on the α value, the minimum temperature to avoid condensation is indicated. This can be seen as the minimum operation temperature for fluidized bed reactors. Temperatures above 350 °C cannot be achieved in fluidized bed reactors due to carbon formation, which limits its application to medium range carbon chain length.

Table 16 Chain growth probability and minimum synthesis temperature to avoid condensation for different product chain lengths[354]

Product number of C atoms	Chain growth probability α	Minimum temperature in °C
2–5	51%	109
5–11	76%	329
5–18	82%	392
12–18	87%	468

Low-Temperature-Fischer–Tropsch (LT-FT) reactors contain liquid hydrocarbons and can hence be operated below condensation temperature. The main products in these reactors are waxes, whereas High-Temperature-Fischer–Tropsch (HT-FT) produces mainly alkenes and gasoline.[36]

Liquid products (in the range of 10–23 carbon atoms per chain) can be used to substitute gasoline or diesel and have advantages compared to conventional fuels, due to their low aromatics content and the absence of sulphur. This results in lower emission values regarding particles, SO_2 and aromatics.

As Fig. 21 shows, depending on the chain growth probability, only a small product fraction can directly be used as a diesel fuel. Other by-products must be upgraded first, which mainly refers to hydro-cracking of waxes.

4.1.2. Catalysts. In commercial processes, Fe- and Co-catalysts are used, whereas nickel and ruthenium-based catalysts, even if catalytically active, play no role in industrial applications. Porous metal oxides with large specific surfaces, for example zeolite and aluminium oxide, are used as catalyst carriers.[355]

Co-based catalysts are mostly used for natural gas as feedstock and have a high activity.[35] Natural gas has a high H_2/CO ratio and no WGS reaction is needed to achieve a suitable syngas composition. Ryttera *et al.* give an overview of different parameters on the selectivity of Co-based catalysts for higher carbons.[356] At lower H_2/CO ratios, less hydrogen is adsorbed on the catalyst surface increasing the chain growth probability. The same is true for lower temperatures, however the CO

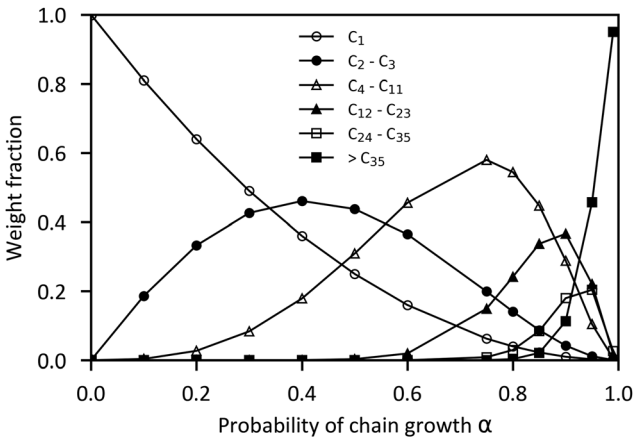


Fig. 20 Anderson–Schulz–Flory distribution: weight fraction of different carbon number products as a function of the chain growth α. α is among other things influenced by the catalyst, the reaction conditions, the temperature and the syngas stoichiometry. Adapted from ref. 351–353.

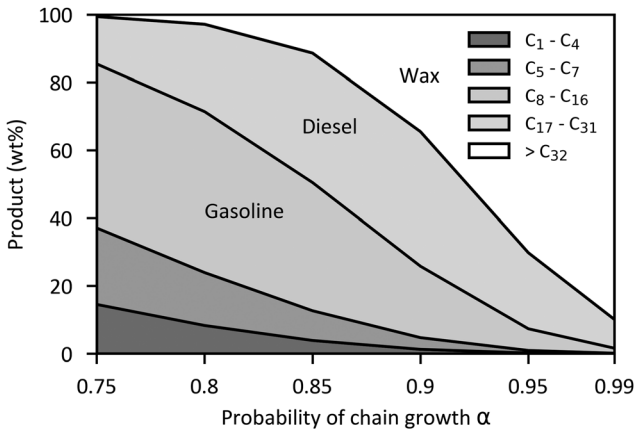


Fig. 21 Fischer–Tropsch product distribution as a function of the chain growth probability, calculated on a mass fraction basis. The number of carbon atoms per chain is shown.

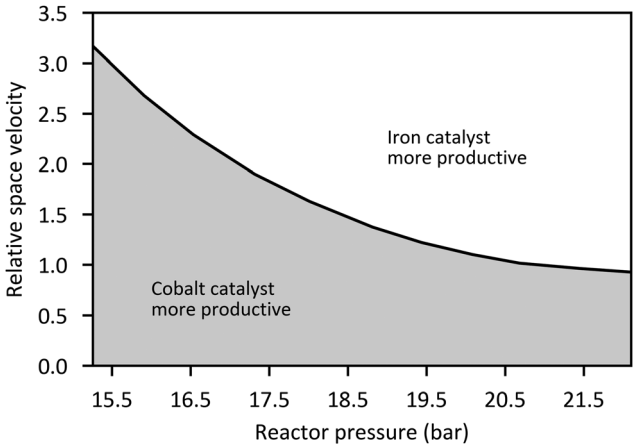


Fig. 22 Productivity comparison between iron (240 °C) and cobalt (220 °C) based catalysts. Adapted from Steynberg and Dry.[35]

conversion decreases.[356] At higher water partial pressures methane formation is inhibited improving selectivity as well.[357,358] The influence of CO conversion (controlled by the GHSV) depends on the reactor type. For slurry reactors selectivity for higher carbons improves until a conversion of around 75% when CO_2 is formed by WGS, which is catalysed due to oxidation of the catalyst.[356] The oxidation is facilitated within the slurry reactor through backmixing. Therefore the same effect does not occur in fixed bed or micro-channel reactors.[356] CO_2 methanation is not possible with Co-catalysts, as Co does not catalyse the WGS reaction and blends of the methanation catalyst with a WGS catalyst are not efficient in CO_2 conversion.[123] For this reason CO_2 is either adsorbed or hydrogenated and therefore only acts as a dilutor.

Compared to Co-catalysts, Fe also catalyses the WGS reaction and is therefore better suited for coal and biomass feedstocks[359] as well as for CO_2 as a feedstock.[123] In this application, higher temperatures above 300 °C are needed due to the higher CO equilibrium concentration. Alkali metals are used as promoters, especially potassium. However their effect highly depends on the interaction with the support material.[360] The best conversion rate was achieved with alkalised Fe on γ-alumina.[361] H_2O on Fe catalysts promotes product inhibition, which leads to a restricted conversion and complicated gas recycling.[359]

Compared to the Fe catalysts, the Co-based catalysts have a higher productivity at higher conversion levels and an equal productivity at intermediate levels.[362] As can be seen in Fig. 22, iron catalysts are favourable at increased pressure and space velocity.

In the first years after discovery (1925–1950), Co-catalysts were used exclusively. Afterwards, Fe-catalysts were developed. The still higher activity and stability of Co-catalysts lead to a revival, *e.g.* in the Sasol slurry-phase distillate process.[123]

In terms of selectivity, two different paths can be realised. In the Sasol-Synthol process, low-molecular-weight products are preferred at high temperatures and all products are in gaseous state at reaction conditions (HTFT). It is operated above 330 °C with Fe-catalysts.[363] The second reaction path takes place at lower temperatures with the highest catalytic activities and produces high-molecular-weight products, which are liquid at reaction conditions (LTFT). Both Fe- and Co-based catalysts are used for this route, which need to have big pore sizes to prevent liquid product plugging.[363] As the product waxes can be hydrocracked to diesel fuels, this process is very flexible. The diesel fuel selectivity of the latter process reaches 80%.[364] One example for a LTFT process is Shell's GtL plant in Malaysia.[365]

The catalysts can be manufactured by impregnation and precipitation. For the former, the metal oxide is impregnated with a metal salt solution and calcinated afterwards.[366] The catalyst activity is influenced by the pore size distribution and the conditions of calcination, but also by promoters *e.g.* alkali metals for Fe-catalysts or potassium and copper for Co-catalysts.

For Co-catalysts, the main causes of catalyst deactivation are poisoning, re-oxidation of Co-active-sites, formation of surface carbon species (*e.g.* coke formation at elevated pressures), carbidisation, surface reconstruction, sintering of Co-crystallites, metal-support solid state reactions and attrition.[367] Two regimes of deactivation can be distinguished: In the initial deactivation regime, during the first days to weeks, reversible deactivation takes place. In the subsequent long-term deactivation regime, irreversible deactivation decreases the syngas conversion rate more slowly but continuously.[367] For Co-catalysts, alkali metals are a catalyst poison. While they increase the chain growth probability, activity is decreased consequently an optimum concentration balancing these effects is necessary.[367] Sulphur is a catalyst poison, especially for Co-catalysts, but also for Fe. Sulphur atoms cause a geometric blockage of the active sites.[367] Therefore, a limit of 1–2 mg m^{-3} or 0.02 mg m^{-3} in modern applications has to be kept.[36,347] The impact of sulphur poisoning depends strongly on the sulphur species.[367]

Fe-catalyst share the same causes of catalyst deactivation, however the oxidation of the active iron phases by water and CO_2 play a more prominent role.[368] Furthermore, deposition of inactive carbonaceous compounds, a loss of surface area by sintering and catalyst poisoning by sulphur are seen as the main deactivation mechanism for Fe-based catalyst.[369]

4.1.3. Commercial converter designs and processes

Converter designs. An overview over four currently used converter designs is given in Table 17. Assigned to the HTFT processes, the Sasol's Synthol (SAS) process has been developed from the Argen process (low temperatures, medium pressures, fixed bed) since the 1950s. It involves higher temperatures (300–600 °C) and medium pressures in a circulating fluidized bed and produces light olefins. The updated version is the Advanced Synthol process.[370]

In Sasol's slurry-phase distillate (SSPD) process, hot syngas bubbles through a slurry of catalyst particles and liquid products at low temperatures (200–250 °C) and pressures. Whereas in the beginning Fe-based catalysts were used, Co-based catalysts are in favour today.[370] The second listed LTFT process is the Shell middle-distillate synthesis (SMDS) process. In this two-step process, paraffinic waxes are synthesised and subsequently hydrocracked to gain a middle-distillate product. In the

Table 17 Overview of commercial FT converter designs

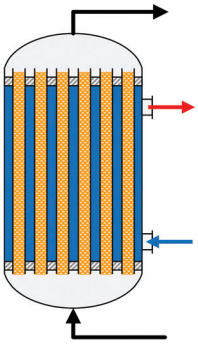
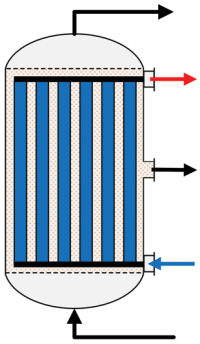
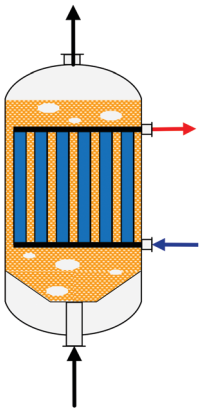
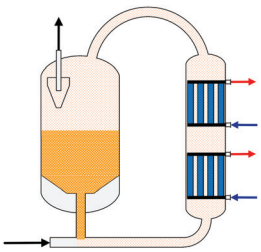
	Fixed bed (trickle bed)	Slurry phase	Fluidized bed	Circulating fluidized bed (CFB)
				
Description	Tubes filled with catalyst, boiling water on the shell side. The products are mostly liquid, therefore the reactor is processed as a trickle bed reactor.	Waxy products are used as heat carrier and suspension fluid for fine catalyst particles. The syngas bubbles through the liquid and cooling coils remove the heat.	Low-chain molecules are produced at high temperatures (>300 °C)	Low-chain molecules are produced at high temperatures (>300 °C)
Temperature limit	<260 °C, higher temperatures lead to catalyst swelling and hence tube blocking	<260 °C, higher temperatures lead to hydrocracking of the waxes	>300 °C, to avoid agglomerating of catalysts by heavy waxes	>300 °C, to avoid agglomerating of catalysts by heavy waxes
Temperature control	Small tube diameters (5 mm for Fe-catalysts, narrower for Co due to activity) and high gas velocities increase the heat flux, medium temperature has to be lower to avoid hotspot temperatures above limit; big temperature gradients in radial and axial direction.	Well mixed reaction phase leads to isothermal conditions and excellent temperature control, therefore average temperature can be higher than in fixed bed. Reactor cooling coils in suspension fluid generating steam	High degrees of turbulence exhibit very high rates of heat exchange; bed is virtually isothermal	High degrees of turbulence exhibit very high rates of heat exchange; bed is virtually isothermal
Process conditions	220–240 °C 20–25 bar	220–240 °C 20–25 bar (SASOL SPD)	320–350 °C Higher pressure up to 40 bar possible because reaction section is much wider	320–350 °C 25 bar
Pressure drop	High in large-scale reactors 4 bar	Lower than fixed bed 1 bar	Much lower than in CFB, due to lower velocity	
Selectivity	Similar product distribution to crude oil (high carbon number products)	Similar product distribution to crude oil (high carbon number products)	Mainly low carbon number products (80% < C_{10}) No condensation on outer side of particles (agglomeration)	Mainly low carbon number products (80% < C_{10}) No condensation on outer side of particles (agglomeration)
Conversion	~60% per pass, recycling is used	~60% per pass; carbon efficiency 75%	>85% (higher than CFB due to more catalyst participation in reaction, which increases the conversion)	~85%
GHSV				3–4 times higher velocities than in FFB
Space charge	Low	Higher than fixed bed		
Construction, operation	High construction costs and high weight for large reactors. Easy separation of liquid products	25% more cost-efficient compared to fixed bed (for Fe-catalysts and same capacity); biggest challenge is separation of wax and catalyst; viscosity limits catalyst loading; Thermal efficiency 60%	Construction cost 40% lower than CFB, support structure 95% cheaper; much smaller Less maintenance than CFB, therefor longer on-line times and higher production rates	Narrower sections are ceramic lined because of abrasive catalyst, regular maintenance is essential
Catalyst	Fe or Co-based	Fe or Co-based	Fe-based	Fe-based
Catalyst loading and consumption	Only catalyst in first centimetres is affected by catalyst poisons	Catalyst loading 4 times lower than in fixed bed as particle size is 1–2 orders of magnitude smaller than in fixed bed while pore diffusion is rate determining Small catalyst loading increases sensibility against catalyst poisons like H_2S	Carbon deposition is less significant, therefore a lower rate of catalyst removal than in CFB is necessary	Carbon deposition results in lower catalyst loading which leads to declining conversion

Table 17 (*continued*)

	Fixed bed (trickle bed)	Slurry phase	Fluidized bed	Circulating fluidized bed (CFB)
Catalyst removal/ exchange	Offline	Online; longer reactor runs, but less advantage with Co-catalysts	Online; due to less carbon deposition, less replacement is necessary than for CFB	Online
Upscaling	Easy by adding tubes, limited to 3000–4500 barrel per d	Easy, up to 20 000 barrels per d		
Example process	Shell Middle Distillate Synthesis (SMDS)	SASOL SPD	SASOL SAS	Synthol
Ref.	123 and 371	35, 123, 307, 371 and 372	35, 36, 371 and 373	35, 36 and 371

heavy paraffin synthesis stage, a tubular fixed bed reactor and a catalyst with high selectivity towards heavier products is used followed by a trickle flow hydrocracking reactor.[370]

Processes. Starting in the 1950s, Sasol built the first commercial, large-scale FT plant for the production of synthetic fuels (Sasol I) in Sasolburg (South Africa). The knowledge gained from operating the first generation of HTFT Circulating Fluidized Bed (CFB) reactors was used to build two more plants (Sasol II and Sasol III) in Secunda in the 1980s, which were based on the second generation of CFB reactors. While Sasol II and III were converted at the end of the 20th century to the Sasol Advanced Synthol reactor type (stationary fluidised bed instead of circulating) which are used until today, Sasol I was completely converted to the production of waxes (LTFT). For this purpose, the CFB reactors were dismantled and a reactor based on the Sasol slurry phase distillate (SPD) process was added to the ARGE Type reactors which had already been operated since the 1950s. The three plants are still in operation today. Whereas in the beginning all synthesis plants used gas from coal gasification, Sasol I is now operated entirely with natural gas. In addition to Sasol, Shell built a comparable high capacity FT plant in the 1990s based on their SMDS fixed bed reactors in Bintulu (Malaysia). Both the SMDS (Shell) and the SPD (Sasol) processes were used by Qatar Petroleum in 2011 and 2007 respectively to install two large-scale GtL plants in Ras Laffan. Sasol itself is currently moving away from coal and natural gas-based synthesis processes towards the production of chemicals.[374] Another HTFT plant focusing on the use of low grade coal is the Shenhua Ningxia CtL plant (China), which was commissioned in 2016. The plant, built with the participation of China Fuels, has a capacity of 100 000 bpd and produces diesel, naphtha and alcohols.[375] An overview of the process parameters of the HTFT and LTFT plants mentioned is given in Tables 18 and 19 respectively.

4.1.4. Product upgrade. If an Fe catalyst is used in LTFT slurry reactors, another downstream gas processing operation is usually combined because it is expensive to achieve high reactant conversions.[35]

High molecular weight products (wax fraction) are either processed to speciality wax products by hydrogenation, hydroisomerisation or controlled oxidation, or converted to naphtha, diesel and fuel jet in a hydro cracker[389–391] Diesel fuel can be gained not only from fractionating, but also from wax hydrocracking and olefin oligomerisation.

Lower molecular weight fractions are hydrogenated and fractionated. The naphtha can be upgraded in a conventional refining step to produce gasoline.[35]

Several distillation and absorption columns separate the resulting products. These processes are well established from petrochemical industry and their influence on the process performance is rather small.[392]

Ethen and propene can be used for polymer production and offer good market prices. Propane and butane can be sold as LPG. The naphtha fraction can be used to produce light olefins.

In a first step, light hydrocarbons and gases are removed to ease atmospheric storage. Then oxygenated compounds are

Table 18 High temperature FT-process examples

	Sasol	Sasol		Mossgas (PetroSA)
Location	Sasolburg	Secunda		Mossel Bay
Plant Name	Sasol I	Sasol II & III (CtL)		
Product		Gasoline, diesel, C_2–C_4 gases		Gasoline, distillates, kerosene, alcohols and LPG
Temperature	340 °C	340 °C		340 °C
Pressure	20 bar	20–40 bar		20–40 bar
Catalyst	Fe-based	Fe-based		Fe-based
Product capacity		160 000 bpd		1 million t per a
Feed	Coal	Coal		NG
Reactor type	CFB 1st generation	CFB 2nd generation	SAS	CFB (Sasol Synthol)
Status	Shut down	Replaced	Running	Running
Time of operation	1950s–1990s	1980s–1990s	Since 1990s	Since 1992
No. of reactors	3	16	10 (formerly 8)	
Ref.	36 and 376–378	376, 377 and 379	36, 376, 377 and 379	35, 36, 376 and 380

Table 19 Low temperature FT-process examples

	Sasol			Shell	Qatar Petroleum, Shell	Qatar Petroleum, Sasol
Location	Sasolburg			Bintulu	Ras Laffan	Ras Laffan
Plant name	Sasol I				Pearl GTL	Oryx GTL
Product	Waxes, alkanes			Waxes	Diesel, naphtha	Diesel, naphtha
Reactor type	ARGE (FB) multitubular	FB (high pressure)	Sasol SPD (slurry)	SMDS fixed bed multitubular	SMDS fixed bed multitubular	Sasol SPD (slurry)
Temperature	230 °C		220–240 °C	200–230 °C	230 °C	230 °C
Pressure	27 bar	45 bar	20–30 bar	30 bar	30 bar	20–30 bar
Status	Running	Running	Running	Running	Running	Running
Catalyst	Fe-based			Co-based	Co-based	Fe-based
Product capacity	2500 bpd (500 bpd each)		2500 bpd	12 500 bpd (8000 bpd each)	140 000 bpd	34 000 bpd
Feed	Coal/NG since 2004	NG	NG	NG		
Time of operation	Since 1950s	1987	Since 1993	Since 1993	Since 2011	Since 2007
No. of reactors	5			4	24	2
Sources	35	381	372, 382 and 383	376 and 384	385 and 386	387 and 388

removed in a liquid–liquid extraction, for example. In the next step, olefins can be removed by fractionating and extractive distillation. For the remaining liquid products and waxes, the conventional petroleum refinery process can be utilised[35,363] even though there are big differences in terms of feed composition, focus of refining and heat management.[393]

4.2. Current research on power-to-FT-fuels

4.2.1. FT-fuels from CO_2. Whereas FT synthesis is well established for NG- and coal-based syngas and several facilities with biomass feedstock exist,[394–397] direct utilisation of CO_2 has not been investigated excessively on a larger scale. However, concepts that recycle CO_2 by feeding it back to the reforming unit (for NG-based FT-synthesis) or directly to the reactor have been investigated. Here it is shifted to CO. As long as the CO/H_2 ratio is adequate, this syngas can be used in the same way as the one from coal gasification or methane reformation.[389] For high recycle ratios above 0.8, both thermal and carbon efficiency are higher for the direct feed of CO_2 to the FT reactor.[389]

For the hydrogenation of CO_2 to alkanes, kinetic equations have been developed.[398,399] As CO_2 is a thermodynamically stable molecule, higher synthesis temperatures are required to improve conversion, which lead to a lower chain growth probability and result in more low-carbon number products.[400] An increase in pressure and a decrease of the gas velocity lead to more saturated products. Furthermore, lower conversion rates as well as higher inert gas and CO_2 content are increasing the chain length.[401,402]

Regarding conventional FT catalysts, Fe catalysts are more promising for CO_2-based feed gases than Co-catalysts because they also catalyse the WGS reaction,[398,400,403–405] whereas the feed gas has to be shifted first for a Co-catalyst. If CO_2 is fed directly to a Co-catalyst, product selectivity is strongly shifted to methane, even at typical FT temperatures. Fe-based catalysts can be used for CO_2 hydrogenation in all three reactor types (fixed bed, fluidized bed, slurry reactor). If the process is aiming for saturated, high boiling products, slurry reactors are advantageous over fixed bed reactor. Fluidized bed reactors are better suited for unsaturated low boiling components. Both slurry and fluidized bed reactors increase the heat transfer from the highly exothermic reaction and are therefore advantageous for the CO_2 conversion and product selectivity.[406] More recently promising Fe-based multifunctional catalysts have been reported, which allow a highly selective conversion of CO_2-based feed gas to gasoline-range hydrocarbons,[407,408] however, tests on a pilot or industrial scale have not been conducted yet.

Due to the limited experiences with CO_2-based feed gases for FT synthesis and the early stage of FT catalyst development for direct CO_2 conversion, realised and planned Power-to-FT-fuels (PtFT-fuels) plants exclusively rely on a shift from CO_2 to CO *via* RWGS-reactor or co-electrolysis. An overview of several projects is given in Table 20.

Sunfire GmbH together with several partners developed a first demonstration plant called 'Fuel 1', which was opened in 2014 in Dresden (Germany). The plant was operated for more than 1500 h and can produce 1 barrel (159 l) of FT product per day. A high-temperature SOEC is used for hydrogen production. CO_2 is converted in a RWGS reactor. The energy efficiency is around 60% and a CO_2 mitigation potential of 85% has been determined compared to conventional fuels.[409–411]

VTT and LUT operated a first demonstrator plant for 300 h in 2017 within the Soletair project. CO_2 was obtained by DAC and H_2 produced by PEM electrolysis. The CO_2 was converted in a RWGS reactor before entering the FT-reactor. A Co-based catalyst was applied. The microstructured heat exchanger reactor was operated between 200–250 °C and 20–30 bar. Due to a mismatch of production capacities, additional CO and H_2 needed to be supplied from gas bundles, also explaining the low carbon efficiency of 38%.[412]

In 2019 a so-called integrated PtL test facility has been opened as part of the 'Kopernikus P2X' initiative funded by the German Federal Ministry of Education and Research. Within a container DAC, co-electrolysis, microstructured FT-reactor and a hydrocracking unit are combined to produce liquid fuels. While the current configuration is only able to

Table 20 Overview of bench, pilot and demo PtL-plants based on FT-synthesis

	Sunfire – fuel 1 plant	Soletair	Integrated PtL test facility	Nordic Blue Crude	The Hague Airport Demo-Plant
Location	Dresden (Germany)	Lappeenranta (Finland)	Karlsruhe (Germany)	Heroya Industripark (Norway)	Rotterdam (Netherlands)
Year	2014	2017	2019	2022	Announced 2019
Project partner	Sunfire, EIFER, Fraunhofer, GETEC,HGM, FZ Jülich, Kerafol, Lufthansa, Univ. Bayreuth, Univ. Stuttgart	VTT, LUT	KIT, Climeworks, Ineratec, Sunfire	Nordic Blue Crude AS, Sunfire, Climeworks, EDL Anlagenbau	Rotterdam The Hague Airport, Climeworks, SkyNRG, EDL, Schiphol, Sunfire, Ineratec, Urban Crossovers
Operating time	>1500 h	300 h		—	—
Fuel production	159 l per d		10 l per d	8000 t per a[a]	1000 l per d[a]
Process					
Carbon source		DAC, 3.8 kg CO_2 per d	DAC	28 000 t per a	DAC
Electrolysis	SOEC, 10 kW	PEM electrolysis, 5 kW, 50 bar	Co-electrolysis (SOEC), 10 kW	Co-electrolysis, 20 MW	Co-electrolysis (SOEC)
Syngas production	RWGS	RWGS-reactor, 850 °C, 1–5 bar	Co-electrolysis (SOEC)		Co-electrolysis
Reactor		Microstructured heat exchanger reactor cooled by high pressure boiling water	Microstructured reactor		Microstructured reactor
Ref.	409–411 and 417	412, 418 and 419	413–415	415 and 416	420

[a] Planned capacity.

supply 10 l per d, a larger pilot plant for the second project phase and a pre-commercial plant design in the third project phase is planned.[413–415]

A first industrial sized PtL-plant based on FT-synthesis is scheduled to open by 2022 in the Heroya industrial park (Norway). The plant will feature a 20 MW co-electrolysis for syngas production and will produce 8000 t per year of FT-crude, which can be further processed in existing refineries. The power for the electrolysis will be supplied by hydropower plants.[415,416]

Another project announced in 2019 is a demo-plant at the Hague airport in Rotterdam. The goal is to produce 1000 l of jet fuel per d. CO_2 will be directly captured from the air and co-electrolysis will be used to produce syngas. A concrete opening date has not been reported yet.

4.2.2. Flexibility and dynamic operation. Usually, FT-synthesis is not conducted in a flexible mode. During the 1980s and 90s, a forced cycling mode was investigated to increase the average reaction rate.[348]

Feimer *et al.* published a series of articles about forced cycling in the 1980s. The first article investigates the response of the reaction rate for different products to feed concentration step changes.[421] It observes that a steady state is reached not less than 20 minutes after changing the hydrogen content in the feed from 1 to 0.65 in the reference case. Whereas in forced cycling this is an aspired effect, it means unsteady product concentrations in regular FT-reactors.

For moderate changes of the H_2/CO ratio (1.1–2), the consumption and formation rates determined under stationary conditions can be used for periodic changes in the H_2/CO ratio, as can the resulting product distribution.[422] Nevertheless, it must be noted that the synthesis of long-chain hydrocarbons follows the changes in process conditions much more slowly than the synthesis of CH_4.[423] However, considering the entire process of FT synthesis itself, clear dependencies on changes in process conditions become apparent with regard to process control (*e.g.* educt separation and recirculation) as well as catalyst activity and product distribution.[422]

A common method to gain a required flexibility in the synthesis is the integration into an overall process. For example, a BtL and PtL based process can achieve flexibility by providing a storage system between the educt supply and the synthesis reactor.[422] Instead of storing excess electricity in H_2 storage facilities after electrolysis, current work indicates that the direct utilisation by adapting the H_2/CO ratio can be used while increasing the productivity and simultaneously keeping the product distribution range almost unchanged.[424]

As already mentioned, the product distribution strongly depends on the contact time between the catalyst and educt gas. Therefore increasing GHSV result in a decrease of the CO conversion as well as the C_{5+}.[425]

4.2.3. Techno-economic analysis. An overview of techno economic studies on FT-based fuel synthesis is given in Table 21. Compared to gasification-based process routes, PtL has higher production costs. Currently, the majority of the available techno-economic studies are focused on biomass to liquid processes.[362,426–428] Analysing the investment costs of

Table 21 Comparison of different techno-economic studies regarding FT-based power-to-liquid and X-to-liquid production routes

	Carbon source	Production cost	Ref.
PtL	Carbon capture	2.3 € per l	344
PtL	Carbon capture	1.2–2.8 $€_{2014}$ per l	432
BtL	Biomass gasification	1.2–2.8 $€_{2014}$ per l	432
BtL	Biomass gasification	1.3 $€_{2012}$ per l	428
CtL	Coal gasification	0.4 $€_{2012}$ per l	428
Diesel	Crude oil	0.4–0.6 $$ per l	433

gasification-based synthesis processes, the share for fuel synthesis is slightly higher than for gasification (67 Mio €/150 MW synthesis *vs.* 77 Mio €/400 MW gasification).[428] In case of currently installed FT plants, 60–70% of the costs are spent for syngas production (coal gasification, methane reforming), 22% for the FT synthesis itself and around 12% for product upgrade and refining.[363,429–431]

Variable H_2 supply costs (electricity costs) and plant capacities given, a change from pure BtL over combined processes to pure PtL can be shown for falling electricity costs and smaller plant capacities as well as decreasing production costs for increasing plant scales.[336,432] The supply of hydrogen in pure PtL processes accounts for the vast majority of the production costs, which also displays the strong coupling to the electricity as shown above.[336,344]

In general, the economic feasibility of conventional as well as pure electricity based FT-plants is strongly connected to the market price of diesel and therefore the crude oil price. The crude oil price represents (as shown in Fig. 23) at least a share of 30–50% of the final diesel pricing. FT can be an economic alternative if some requirements are fulfilled. First of all the carbon source must be large enough and cheap, for example stranded or associated natural gas, which is mostly flared now, or to be economically friendly based on carbon capture or biomass gasification. For the history of the FT process, the development of crude oil prices was crucial. In times of an increasing oil price, like in the 1940s and during the oil crisis of the 1970s, FT plants were commissioned.[36] At least a product market, *e.g.* fuel-driven cars, must be accessible. If currently expected PtL prices are taken as a basis, it can be seen that even if FT-based diesel based solely on electricity were to be fully released from taxes, it would not be able to compete in Germany. In order to make FT-diesel marketable, an additional charge on fossil fuels would therefore have to be raised in addition to the tax release for sustainable fuels.

In addition to pure techno-economic approaches, Tremel *et al.* used a ranking system, which includes acceptance as a third group of indicators, to compare different H_2 based fuels. The ranking system contains about 20 different weighted indicators, including available infrastructure, number of product upgrading stages or market prices. Compared to SNG, ammonia, DME and methanol, FT-fuels achieved the second highest ranking in the final evaluation.[78] This shows in particular that, mainly in a phase-out period in which vehicles still powered by internal combustion engines are used, FT-fuels can make a significant contribution to reducing the demand for fossil fuels. In conclusion, FT-diesel would be economically usable if the price lies below the crude oil price.

Neglecting H_2-production coupled costs the synthesis capital costs can be indicated as one of the next most costly factors, whereas the CO_2 capture or water costs can be neglected.[336]

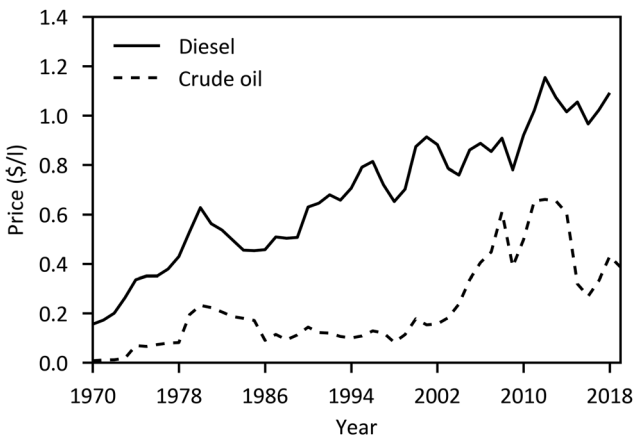


Fig. 23 Development of crude oil and diesel prices in Germany over time, own illustration based on ref. 434 and 435.

4.3. Discussion

Fischer–Tropsch synthesis is a well-known process that has been developed for almost a century. Big plants with coal and natural gas as a feedstock exist and can be economic, depending on the framework conditions. Smaller plants with biomass-based feedstock have been built in the last decade. For PtFT-fuels plant concepts an opposing trend compared to state of the art FT-processes can be observed. While conventional FT plants have increased in scale to achieve profitability by economy-of-scale effects, many realised PtFT-fuels pilot plants are modular, small-scale, all-in-one container concepts. Their goal is to become profitable by allowing a more dezentrailzed, flexible operation and an easy scale up through modularity.

There is still need for research on FT synthesis based on CO_2. On the one hand, direct FT-fuel synthesis from CO_2-based feed gas is still at a very early stage, requiring further catalyst developments and first lab scale plant tests before becoming an option for future PtL plants. On the other hand, several PtFT-fuels demo plants that include a shift from CO_2 to CO have been operated successfully and further larger-scale plants have been announced. For the near term future this will remain the dominant process design for FT-based PtL plants. Consequently, FT-based PtL can benefit substantially from developments in RWGS reactor design and co-electrolysis technologies.

FT-synthesis has not been operated in a flexible mode at a larger scale and research on flexibility and dynamic operation is limited. As a result first PtFT-fuels pilot plants have usually been operated continuously and large hydrogen as well as CO_2 storages seem to be currently unavoidable for large-scale PtFT-fuels plants. Further research is necessary to determine the influence of pressure and temperature changes, variations in the syngas quality and part-load performance to determine an operation window in which a satisfying amount of liquid fuels can still be produced. In addition, plant and reactor concepts that allow a dynamic switch between part load and full load operation should be investigated to be able to benefit from low electricity prices without high storage costs. The current state of research does not yet permit cost-optimised or competitive operation of PtFT-fuels plants. Further techno-economic analyses must be undertaken in order to identify possible markets or routes to market entrance and to guide the research accordingly.

5. Remarks on future power-to-liquid applications

Identifying the optimal PtL process route and selecting the most promising product is still challenging and general recommendations remain difficult because the final decision depends on numerous factors including technological, political and economical aspects. Nonetheless some general aspects that should be considered can be derived.

Regarding water electrolysis technologies, currently AEL offers the highest technology readiness. In the short-term therefore it can be seen as the obvious choice for relatively low-cost, reliable electrolysis technology for PtL, in particular for large-scale plants. As improvements in the technology and cost reductions have been rapid for PEMEL and with its higher flexibility, it seems especially suited for PtL processes providing grid stability through flexible operation. In the long-term SOEC is very promising for PtL due to the ability to use the process heat generated by the exothermal synthesis reactions thereby achieving higher overall process efficiencies. This will be particularly beneficial in cases where the PtL plant is not integrated into existing industrial processes that can use the excess heat and by-products. Furthermore, the possible application of SOEC for co-electrolysis to produce syngas is favorable for PtFT-fuels and the direct PtDME synthesis route. However, the technology readiness still lacks behind and pressurized operation remains difficult. Therefore, it is a confirmation of the advantages SOEC has to offer for PtL that some of the large-scale projects that have been announced are planning to use co-electrolysis in a MW-scale in only a few years (see Section 4.2.1).

The provision of hydrogen *via* water electrolysis is the major cost factor for all PtL products. Therefore reducing the hydrogen cost is a significant challenge for the commercialisation of PtL processes. However, depending on the PtL product, different cost reductions are necessary to achieve cost competitiveness. Based on the evaluations by Schemme *et al.*[344] cost shares of the H_2 supply for the discussed PtL process routes can be estimated. In the case of methanol synthesis, the H_2 supply covers a share of approximately 83%. Assuming that the continuous development of electrolysis and falling electricity prices are the greatest potential for cost reduction, a cost decrease of 25% for the H_2 supply may already be sufficient to make PtMeOH economically competitive. Similar to the PtMeOH process, the PtDME process with a H_2 supply share of 81% would need a substantial reduction of the H_2 supply costs. In contrast, PtFT-fuels would not be viable as a Diesel substitute even with free hydrogen supply due to the low crude oil prices. This, however, may not be the case if FT-waxes are considered as main synthesis product. The change from fuel-focused Coal- or Gas-to-Liquids to the production of chemical products was also announced in 2017 by Sasol's joint chief as a new corporate strategy.[374] Especially in combination with existing FT plants, modifying reactors towards wax production and a change to CO_2/H_2-based feedgas could lead to an economically viable introduction of PtFT-waxes processes.

Currently and in the medium-term CO_2 point sources that offer a high partial pressure of CO_2 in their flue gases remain the most cost-efficient carbon source for PtL (see Section 1.1.2). Due to the high capital investment necessary to build a PtL plant, however, security of the CO_2-supply must be considered. This can be critical for PtL as CCU technology integrated into existing lignite, coal or gas plants that might be shut down in government efforts to achieve climate targets. As a consequence, industrial CO_2-sources *e.g.* from the cement or steel industry as well as CO_2 from biogas plants are specifically interesting for PtL. In the long-term, with sufficient cost reductions DAC can provide a certain geographic independence. This might be particularly important for small-scale PtL plants designed to provide grid flexibility as well as large-scale plants build close to large solar or wind power plants away from industrial CO_2-sources.

Regarding the purity of the CO_2-source, all considered syntheses require very low concentrations of sulphur and chloride compounds making gas cleaning necessary for most carbon sources. PtMeOH offers the advantage that CO_2-based feedgases can be used directly and most catalysts are able to tolerate CO as well as CO_2 allowing for fluctuations in the feed. The same is true for two-stage PtDME processes. In contrast PtFT-fuels currently requires pure H_2/CO syngas with low CO_2 content as feed.

Regarding the scale of PtL plants, opposing trends can be observed. Many PtMeOH and PtFT-fuels research projects are focusing on small-scale, modular container concepts. First commercial projects *e.g.* George Olah plant or Nordic Blue Crude, however, rely on large-scale processes. This illustrates that the scale of future PtL plants is still up for discussion and will probably depend on the surrounding factors of the project as well as the purpose of the plant. Commercial PtL processes trying to provide a carbon-neutral substitute to fossil fuel based products might require a large scale to become cost competitive and producing the required volumes. PtL as a sector coupling technology providing flexibility to the power grid, however, can benefit from small-scale plants that can be implemented close to wind or solar power plants and have the flexibility to use excess electricity. Under this premise, large-scale PtL concepts should focus on reducing costs, enhancing reliability and maximising operation hours, while small-scale, modular concepts also need to consider flexibility.

The location of the project will significantly influence which product is more desirable. FT-fuels might be particularly interesting for locations were the infrastructure for oil processing is already in place. Hence, for many sun-rich oil producing countries, PtFT-fuels can be considered as a promising technology for the transition away from fossil fuels. Large-scale solar power plants can provide affordable, carbon-neutral electricity and during the transition flue gases from refineries and natural gas processing can be used as CO_2-sources. In the long-term these CO_2-sources can be substituted by DAC to achieve carbon neutrality and allow the production independent from other industrial CO_2-sources. As a consequence PtFT-fuels might integrate especially well into the current economic model of these countries. PtMeOH and PtDME however could become more interesting close to their target market. Methanol already is an important platform chemical, which makes an integration into

existing industrial parks particularly interesting, saving transportation costs and allowing an optimal by-product and heat integration. The same could be true for DME which has been popular in China were it has a large market as a LPG substitute in cosmetics, lighters or as transportation fuel (see Section 3.2.3).

6. Conclusion

PtL is a promising technical solution for the defossilisation of hard-to-abate sectors. It enables sector coupling by carbon capture and utilisation and can be used to produce valuable products for the chemical industry. With regard to the synthesis step, major challenges are the direct utilisation of CO_2, flexible and dynamic operation as well as techno-economic guidance to improve design choices. To which extent these challenges have been addressed depends on the PtL product.

Direct synthesis of methanol from CO_2-based feed gas has already been tested in PtMeOH pilot plants and even on a commercial scale. For DME direct and two-stage synthesis from CO_2 have been investigated, however, pilot plants are still missing and the optimal process route for PtDME remains to be identified. Finally, several PtFT-fuels demo plants based on FT-synthesis have been implemented but all include a shift from CO_2 to CO. Direct utilisation of CO_2 in the FT-synthesis is at an early stage and additional catalyst development as well as lab-scale testing is necessary.

As state-of-the-art methanol, DME and FT synthesis are typically operated at full-load and part-load operation has not been a focus historically, research on flexibility and dynamic synthesis behaviour is generally limited. The economic viability of PtL processes, however, will depend to a large extent on a more dynamic and flexible operation that enables the optimal use of excess power and low electricity prices. Consequently, additional research concerning the catalyst behaviour under flexible conditions, the influence of reactor and process design choices on the dynamics and the determination of the dynamic plant behaviour should be given more attention. In addition, the scope of future analyses should be extended to the entire system and thus the individual subsystems should be examined from a system integration perspective. Among others, this includes the use of by-products, heat integration and synergy effects.

With regards to techno-economic analyses, several studies have been conducted for each PtL product. Independent of the product, several studies have identified the supply of hydrogen and thus both electrolysis and electricity costs as the main cost factors. Likewise, the CO_2-capture is the second major part of the manufacturing costs. Nonetheless detailed cost estimates of the processes are important to support future design choices. This is often made difficult through the extreme variation in boundary conditions. Future studies should therefore strive for more consistent boundary conditions and indicate the influence of changes in their economic assumptions.

Existing pilot plants and future concepts indicate that the development of synthesis concepts for PtL is turning away from the conventional approach in the chemical industry: striving for ever larger-scale plants operated centrally in industry parks. Instead more modular, small-scale concepts dominate, that allow a decentralised operation close to renewable energy sources and CO_2-sources. This trend is particularly pronounced for FT-based PtL concepts, where several container-based concepts have been tested and smaller, microchannel reactors replace the large-scale, commercial fixed bed or slurry reactors. Nevertheless, large-scale projects are still being planned, which leads to a split in the development trend. On the one hand, small concepts in container size, which are designed for decentralized application, can be assigned to the electricity sector. The ideas are usually based on a highly flexible operation to use low electricity prices and to act as operating reserve. The large-scale plants, on the other hand, can be allocated to the chemical or mobility sector and aim for a substitution of current energy carriers and basic chemicals. Within these trends, further investigations into possible business cases must be carried out.

Overall, the commercialisation of PtL will require further advancements of the product synthesis together with improvements in electrolysis and carbon capture.

Conflicts of interest

There are no conflicts to declare.

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