



Ambiphilicity of a mononuclear cobalt(III) superoxo complex†

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Addition of HOTf to a mixture of Co^{III}(BDPP)(O₂•) (1, H₂BDPP = 2,6-bis((2-(S)-diphenylhydroxymethyl-1-pyrrolidinyl)methyl)pyridine) and Cp*₂Fe produced H₂O₂ in high yield implying formation of Co^{III}(BDPP)(OOH) (3), and reaction of Sc(OTf)₃ with the same mixture gave a peroxo-bridged Co^{III}/Sc^{III} 5. These findings demonstrate the ambiphilic property of Co^{III}-superoxo 1.

Metal-superoxo species are often believed to be the first intermediate following dioxygen (O₂) association in the catalytic cycle of O₂ activating metalloenzymes.¹ Despite intensive work in the past, the chemistry of metal-superoxo complexes remains largely unexplored, and hence attracts significant attention from chemists and biochemists.² *Inter alia*, metal-superoxo intermediates can react with NO• or organic radicals to furnish metal-peroxynitrite³ and -alkylperoxo⁴ complexes *via* radical coupling. Furthermore, they exhibit considerable electrophilicity as indicated by their capability of performing hydrogen atom abstraction (HAA)⁵ from weak C–H and O–H bonds and oxygen atom transfer⁶ to triphenylphosphine or thiol anisoles. On the other hand, they can initiate deformylation processes when treated with 2-phenylpropionaldehyde, thereby revealing their nucleophilic character.^{7,8} Besides the aforementioned well know activities, in a given elementary transformation metal-superoxo intermediates may function not only as an electrophile but also as a nucleophile. In fact, ambiphilicity

of metal-superoxo species has been postulated in a series of theoretical and experimental investigations including O₂ activation catalyzed by α-ketoglutarate dependent dioxygenases,⁹ and by Cu, Fe and Co model complexes.¹⁰ Only recently has such ambiphilic property been experimentally confirmed.⁸ In our continuing efforts devoted to investigating reactivity of metal-superoxo intermediates, some of us succeeded in preparing a range of homologous Fe^{III}-, Co^{III}- and Mn^{III}-superoxo species by reacting O₂ with the corresponding divalent precursors.¹¹ It has been shown that these trivalent metal-superoxo complexes can convert into the metal-hydroperoxo complexes *via* HAA. In particular, the reaction of Mn^{III}(BDP^{BPr}P)(O₂•) (H₂BDP^{BPr}P = 2,6-bis((2-(S)-di(4-bromo)-phenylhydroxymethyl-1-pyrrolidinyl)methyl)pyridine) with trifluoroacetic acid (TFA) and Sc(OTf)₃ yields rare examples of Mn^{IV}-hydroperoxo complexes, Mn^{IV}(BDP^{BPr}P)(OOH), and [Mn^{IV}(μ-OO)Sc(OTf)_n]⁽³⁻ⁿ⁾⁺ as evidenced by the combined spectroscopic and computational studies (Scheme 1).⁸ Obviously, these proton- and metal-coupled electron transfer processes provide the first experimental support for the proposed ambiphilicity of metal-superoxo species. In this regard, more examples are desired to fully understand how the ambiphilic property of metal-superoxo species affects their chemical reactivity. To this end, we examined the reaction of a Co^{III}-superoxo complex, Co^{III}(BDPP)(O₂•) (1, H₂BDPP = 2,6-bis((2-(S)-diphenylhydroxymethyl-1-pyrrolidinyl)methyl)pyridine) with TFA and Sc(OTf)₃ together with external electron donors.

Treating 1 with HOTf in THF at –90 °C gave a gray-green solution attributed to intermediate 2 having two weak absorption bands at 470 and 640 nm, which reached maxima when 1 equiv. of HOTf was added (the inset of Fig. 1). The existence of an isosbestic point at 590 suggested that no intermediate was formed in the course of conversion of 1 to 2 (Fig. 1). Conversely, complex 1 can be retrieved from deprotonation of 2 by 1 equiv. of 1,8-diazabicyclo[5.4.0]undec-7-ene (DBU) with a yield of 80% with respect to 1 (Fig. S1, ESI†). Moreover, complex 2 can be obtained from one-electron oxidation of the hydroperoxo complex Co^{III}(BDPP)(OOH) (3). Adding equimolar of tris(4-bromophenyl)ammoniumyl hexachloroantimonate, which is often referred to as magic blue, to a THF solution of 3 at –90 °C resulted in a gray-green solution, whose

^a Department of Chemistry, National Taiwan Normal University, Taipei 11677, Taiwan. E-mail: wzlee@ntnu.edu.tw

^b Universität Göttingen, Institut für Anorganische Chemie, D-37077 Göttingen, Germany. E-mail: franc.meyer@chemie.uni-goettingen.de

^c Max-Planck-Institut für Chemische Energiekonversion, Mülheim an der Ruhr D-45470, Germany. E-mail: eckhard.bill@cec.mpg.de

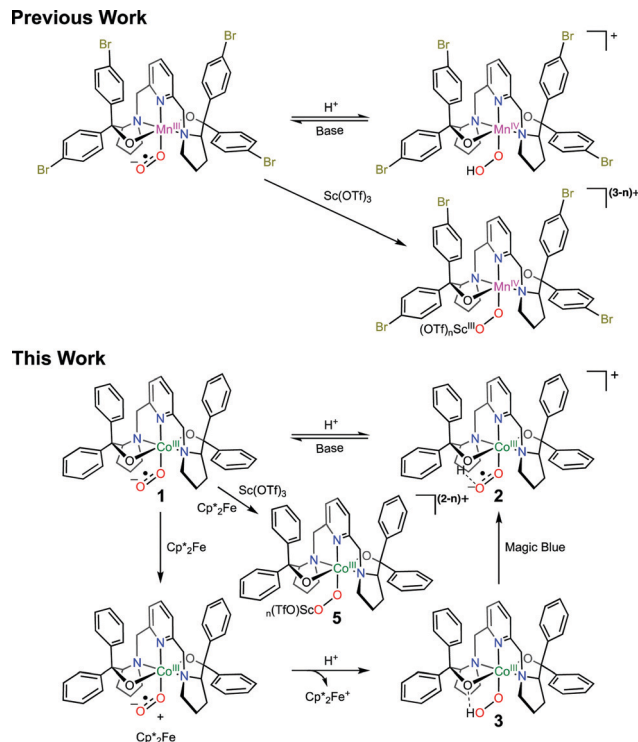
^d State Key Laboratory of Catalysis, Dalian Institute of Chemical Physics, Chinese Academy of Sciences, Dalian 116023, China. E-mail: shengfa.ye@dicp.ac.cn

^e Max-Planck-Institut für Kohlenforschung, Mülheim an der Ruhr, D-45470, Germany. E-mail: shengfa.ye@kofo.mpg.de

^f Department of Medicinal and Applied Chemistry, Kaohsiung Medical University, Kaohsiung 807, Taiwan

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Scheme 1 Ambiphilicity of metal-superoxo complexes.

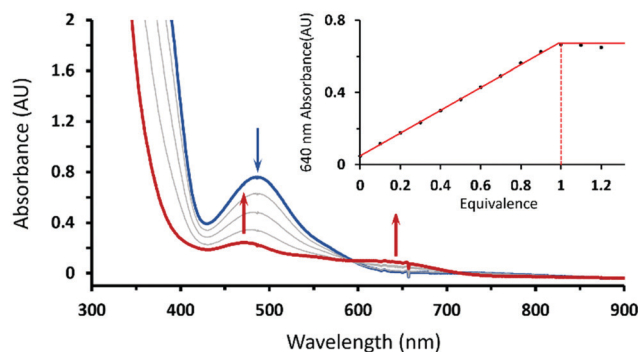


Fig. 1 UV-vis spectral changes of the reaction of **1** (1.0 mM, blue trace) with 1 equiv. of HOTf in THF at -90°C . Inset: Titration curve of **1** with HOTf. X-axis: the equivalents of HOTf, Y-axis: the conversion ratio of the absorption peak at 640 nm.

absorption spectrum displayed the same signature features as those found for **2** (Fig. S2, ESI[†]).

To identify the exact nature of the resulting species **2**, we have undertaken detailed spectroscopic characterization and DFT calculations. The EPR measurement of **2** exhibited a spectrum similar to that of **1** ($A_{\text{Co}} = 18\text{ G}$) except for a slightly larger ^{59}Co hyperfine coupling constant ($A_{\text{Co}} = 24\text{ G}$) seen in Fig. 2, thus indicating that **2** still consists of a Co^{III} center coupled with a radical ligand yielding an overall doublet ground state. The radical ligand thus would be a hydroperoxyl radical or a superoxo having a strong hydrogen bonding interaction with the protonated BDPP^{2-} ligand (Scheme 1) as



Fig. 2 X-band EPR spectra of (A) **1** ($g_1 = 2.098$, $A_{\text{Co}} = 18\text{ G}$); (B) **1** with 1 equiv. of HOTf ($g_1 = 2.066$, $A_{\text{Co}} = 24\text{ G}$); (C) **1** with 1 equiv. of DOTf ($g_1 = 2.064$, $A_{\text{Co}} = 24\text{ G}$) and (D) **1** with 1 equiv. of $\text{Sc}(\text{OTf})_3$ ($g_1 = 2.066$, $A_{\text{Co}} = 24\text{ G}$). Measurement condition: $T = 77\text{ K}$, $f_{\text{mw}} = 9.6\text{ GHz}$.

suggested by the crystal structure of **3**.^{11b} However, the EPR spectrum of the product generated by reacting **1** with deuterated triflate acid (DOTf) is almost identical to that of **2** without discernable line broadening, which essentially rules out the possibility of the radical ligand being a hydroperoxyl radical (Fig. 2C). Repeated attempts to obtain the O–O vibrational frequencies of **2** from resonance Raman measurements did not accomplish, largely because **2** has only weak chromophores in the usual UV-vis region (Fig. 1). Consequently, the intensity of the O–O stretching signal is too low to be readily detected.

DFT calculations also suggested the O donor of the BDPP^{2-} ligand to be the favored protonation site of **1**, consistent with experiment. Even when the starting geometry contained a OOH ligand in which the distal H atom forms a hydrogen bond with the BDPP^{2-} ligand, the geometry optimizations invariably shifted the H atom back to the O atom of BDPP^{2-} and eventually converged to **A** (Fig. 3). We also tested the initial geometry without the hydrogen bond by tilting the H atom upward. The computations indeed yielded a Co^{III} center bound to a hydroperoxyl radical ligand (**C**), but **C** lies $23.2\text{ kcal mol}^{-1}$ higher in energy above **A** (Fig. 3). Moreover, formation of a hydrogen bond between the superoxo motif and the proton of the OH group of the protonated BDPP^{2-} ligand stabilized **A** by $10.4\text{ kcal mol}^{-1}$ relative to **B** (Fig. 3). Thus, **A** is best deemed as the most appropriate model for **2**.

Alternatively, to transform **1** into the corresponding peroxy product, we then added 1 equiv. of decamethylferrocene (Cp^*_2Fe) or sodium naphthalenide ($\text{NaC}_{10}\text{H}_8$) to THF solutions of **1** at -90°C , but UV-vis measurements suggested that no reactions occurred (Fig. S3 and S4, ESI[†]). Taken together, neither proton nor electron donors alone can realize the superoxo-to-peroxy conversion for **1**.

Interestingly, upon treating a mixture containing equimolar **1** and Cp^*_2Fe with 1 equiv. of HOTf, the color of the reaction solution gradually changed from gray-green to dark green then orange; meanwhile, characteristic features of decamethylferrocenium (Cp^*_2Fe^+) emerged suggesting that Co^{III} -superoxo **1** was reduced in the presence of both HOTf and Cp^*_2Fe (Fig. 4A).





Fig. 3 Computed models for **2**.

During this process, we did not observe the formation of Co^{III}-hydroperoxo **3**. Instead, the reaction produced 19% of H₂O₂ with respect to **1**, as determined by iodometric titration (Fig. S5, ESI[†]). When 2 equiv. of HOTf was added, 42% of H₂O₂ was furnished (Fig. S6, ESI[†]). Thus, we reasoned that the aforementioned reaction indeed generates **3**; however, once formed, **3** further reacted with HOTf to produce H₂O₂. On the other hand,

treating **2** with 1 equiv. of Cp*₂Fe (Fig. 4B) also generated 23% of H₂O₂ (Fig. S7, ESI[†]). Therefore, all experimental findings revealed that transformation of **1** to **3** proceeds *via* concerted proton coupled electron transfer, which clearly demonstrated the amphiphilicity of **1**.

Reaction of **1** with Sc(OTf)₃ in THF at -90 °C generated a product, which showed nearly identical absorption and EPR spectra to those of **2** (Fig. S8 (ESI[†]) and Fig. 2D). On the basis of these observations, we tentatively suggest that the reaction furnished a Co^{III}-superoxo···Sc^{III} species, [Co(BDPP)(O₂·)···Sc(OTf)_n]^{[3-n]+} (**4**). Furthermore, addition of Sc(OTf)₃ to the mixture of **1** and Cp*₂Fe in THF at -90 °C afforded a Co^{III}-peroxo-Sc^{III} complex, [Co(BDPP)(μ-OO)Sc(OTf)_n]^{[2-n]+} (**5**), as depicted in Scheme 1 (Fig. S9, ESI[†]). Addition of 2 equiv. of HOTf to the THF solution of **5** produced 76% H₂O₂ as quantified by iodometric titration experiments (Fig. S10, ESI[†]). Formation of H₂O₂ thus strongly supports the identity of peroxo-bridged binuclear **5**. These findings further reinforce the amphiphilic property of **1**.

In comparison with the similar reaction found for Mn^{III}(BDP^{Bt}P)(O₂·) (Scheme 1), the difference can be readily attributed to the much higher oxidation potential of Co^{III} to Co^{IV} than that of Mn^{III} to Mn^{IV}, which can ultimately be rooted back to the distinct effective nuclear charge of low spin Co^{III} compared to high spin Mn^{III} centers. Therefore, formation of an otherwise hydroperoxo O-H bond does not provide a sufficient driving force to trigger an electron transfer from the Co^{III} center to the superoxo ligand. Consequently, the superoxo motif is not electron rich enough to accommodate the incoming proton from HOTf and protonation of the supporting BDPP²⁻ ligand is preferred. Therefore, to effect surperoxo-to-peroxo conversion for **1**, an external electron source has to be provided in addition to Brønsted or Lewis acids.

In conclusion, treatment of Co^{III}-superoxo **1** with HOTf and Sc(OTf)₃ afforded the ligand-protonated Co^{III}-superoxo **2** with a hydrogen bond formed between the O₂·⁻ motif and the protonated BDPP²⁻ ligand and a superoxo-bridged binuclear Co^{III}/Sc^{III} **4**, and Co^{III}-superoxo **1** can be regenerated from deprotonation of **2** by DBU. However, addition of 2 equiv. of HOTf into the reaction mixture of **1** and Cp*₂Fe produced 42% of H₂O₂ suggesting the formation of Co^{III}-hydroperoxo **3**, and the reaction of Sc(OTf)₃ with **1** in the presence of Cp*₂Fe gave a peroxo-bridged binuclear Co^{III}/Sc^{III} **5**. These findings provided strong experimental support for the amphiphilic property of Co^{III}-superoxo **1**. Interestingly, the ligand-protonated Co^{III}-superoxo **2** can be prepared from one-electron oxidation of Co^{III}-hydroperoxo **3**. The unveiled results underline the critical property of amphiphilicity for metal-superoxo species and direct us to design further investigation strategies towards better understanding O₂ activation processes carried out by metalloenzymes and related catalysts.

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Fig. 4 UV-vis spectral changes of addition of (A) 1 equiv. of HOTf into a solution of **1** (1.0 mM) in the presence of 1 equiv. of Cp*₂Fe and (B) 1 equiv. of Cp*₂Fe into a solution of **2** (1.0 mM) in THF at -90 °C.



Conflicts of interest

There are no conflicts to declare.

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