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## RESEARCH ARTICLE

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## Exploring the microscopic mechanism of pseudocapacitance with electronic structures in monolayer 1T-MoS<sub>2</sub> electrodes for supercapacitors†

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Quasi-two-dimensional 1T-MoS<sub>2</sub> is a promising pseudocapacitance ( $C_{\text{redox}}$ ) electrode material due to its large specific surface area, superior electrical conductivity and mechanical stability. However, the microscopic mechanism of  $C_{\text{redox}}$  and its further manipulation *via* modulating the structures and electronic structures are still unclear. Thus, the  $C_{\text{redox}}$  of monolayer 1T-MoS<sub>2</sub> has been explored based on first-principles calculations. For monolayer 1T-MoS<sub>2</sub> adsorbed by H<sup>+</sup> ions on one side or both sides, a band gap opens, decreases and even disappears with the coverage increase in H<sup>+</sup> ions up to 100%. In this process, the charge transfer from the monolayer 1T-MoS<sub>2</sub> to the adsorbed H<sup>+</sup> ions almost linearly increases with the coverage increase in the H<sup>+</sup> ions. In contrast, the potential change rate of the monolayer 1T-MoS<sub>2</sub> reduces, resulting in the enhancement of  $C_{\text{redox}}$ . Herein, the maximum values of  $C_{\text{redox}}$  reached  $\sim 76.7 \mu\text{F cm}^{-2}$  ( $252.8 \text{ F g}^{-1}$ ) and  $\sim 213.7 \mu\text{F cm}^{-2}$  ( $704.5 \text{ F g}^{-1}$ ) for the 100% coverage of H<sup>+</sup> ions on one side and both sides, respectively. Furthermore, the manipulation of  $C_{\text{redox}}$  in the monolayer 1T-MoS<sub>2</sub> could be realized through intrinsic defects engineering. Particularly, the  $C_{\text{redox}}$  was greatly improved in the system with S vacancies. These results would provide significantly fundamental insights for understanding the correlation between the  $C_{\text{redox}}$  and the electronic structures of 1T-MoS<sub>2</sub> and other similar quasi-two dimensional materials.

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## Introduction

With the exhaustion of fossil resources and serious environmental pollution issues, exploiting environmentally friendly and renewable resources is an effective strategy to solve this problem. Among these solving schemes, it is significant to design highly efficient energy conversion and storage systems.<sup>1–6</sup> Nowadays, lithium-ion batteries and supercapacitors are two types of energy storage systems widely applied in electrical/hybrid vehicles and portable devices. Lithium-ion batteries can provide large energy density, *e.g.* 120–170 W h kg<sup>−1</sup>, due to their high redox potential of Faraday

energy storage mechanism. Nevertheless, lithium-ion batteries cannot provide huge energy in a short time to meet the demands of high power devices due to their slower kinetic Faraday process.<sup>7</sup> In contrast, supercapacitors have been extensively studied due to their high power density, fast charging–discharging rate, long cyclic lifetime, low maintenance cost, *etc.*<sup>8–10</sup> However, the current energy density of the supercapacitors is only limited to 5–10 W h kg<sup>−1</sup>.<sup>11</sup> Therefore, it is required and currently urgent to improve the energy density of supercapacitors while maintaining their large power density output.

In order to enhance the energy density of the supercapacitors, improving the specific capacitance of electrodes is one of the important routes. According to the energy storage mechanisms, the supercapacitors are generally divided into two types, namely electrical double layer capacitance (EDLC) and pseudocapacitance ( $C_{\text{redox}}$ ). Remarkably, it is pointed out that the  $C_{\text{redox}}$  is one to two orders of magnitude higher than EDLCs in most of the present supercapacitors.<sup>12</sup> As suitable electrode materials for supercapacitors, it is expected to possess large specific surface area, excellent electrical conductivity, mechanical stability and highly redox reaction activity. Since the discovery of graphene,<sup>13</sup>

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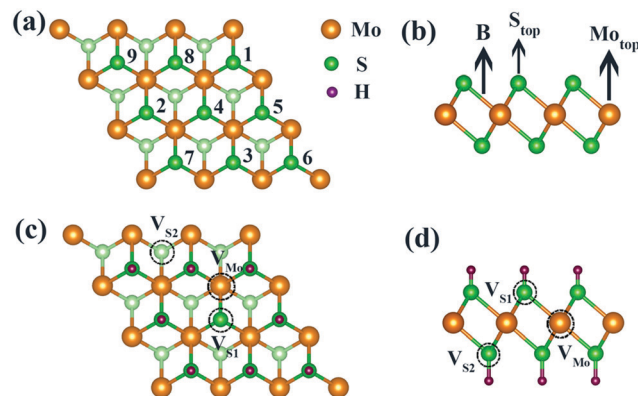
graphene and graphene-like quasi-two-dimensional materials, such as transition metal dichalcogenides (TMDs), graphdiyne, transition metal carbides and carbonitrides (MXenes) have been widely used in electrode materials and photoelectrocatalysts due to their large specific surface area, superior electrical conductivity and good mechanical properties.<sup>14–18</sup> It is noted that the  $C_{\text{redox}}$  or electrochemical performance of graphene and its derivatives could be improved only after introducing intrinsic defects and external dopings, or by combining with transition metal compounds.<sup>19–24</sup> The  $C_{\text{redox}}$  of two-dimensional metal carbides or nitrides is also influenced by surface functional groups dependent on experimental conditions of etching preparations.<sup>25–28</sup>

Monolayer or few layers TMDs, represented by  $\text{MoS}_2$ , have similar structures to graphene and can be well-prepared in experiments. Particularly,  $\text{MoS}_2$  with 1T structure is metallic and it has been applied as electrode materials. For instance, Zhu *et al.* demonstrated that the capacitance of multilayer 1T- $\text{MoS}_2$  in aqueous condition is up to  $380 \text{ F g}^{-1}$  at a scan rate of  $5 \text{ mV s}^{-1}$ .<sup>29</sup> Chhowalla *et al.* displayed that the capacitance of exfoliated 1T- $\text{MoS}_2$  nanosheets ranges from  $400 \text{ F cm}^{-3}$  to  $700 \text{ F cm}^{-3}$  in aqueous electrolytes.<sup>30</sup> In fact, the adsorption of the ions on electrodes not only changes the charge filling of the energy bands but also influences the electronic structures of the electrodes.<sup>31,32</sup> Although considerable experimental progress has been achieved for 1T- $\text{MoS}_2$  as electrodes, the microscopic mechanism of the  $C_{\text{redox}}$  and the manipulation of  $C_{\text{redox}}$  via modulating the structures and electronic structures are still inaccessible.

In this study, the microscopic mechanism of  $C_{\text{redox}}$  with the electronic structures was investigated. We found that the charge transfer almost linearly increased with the increase in the  $\text{H}^+$  ions coverage on one side and both sides of the monolayer 1T- $\text{MoS}_2$ . The  $C_{\text{redox}}$  was dependent on the potential change to a large extent during the process of the  $\text{H}^+$  ions adsorption. The maximum  $C_{\text{redox}}$  was achieved at the 100%  $\text{H}^+$  coverage and further improvements could be realized by introduction of the intrinsic S vacancies. These results clarified the correlation between the  $C_{\text{redox}}$  and the electronic structures of monolayer 1T- $\text{MoS}_2$  and would also provide fundamental insights for the further manipulation of  $C_{\text{redox}}$  in 1T- $\text{MoS}_2$  and other similar 2D materials.

## Computational model and details

Our calculations were performed within the Vienna ab initio simulation package (VASP) based on the plane wave basis method.<sup>33</sup> For the exchange correlation potential, generalized gradient approximation (GGA) with Perdew–Burke–Ernzerhof (PBE) functional was adopted.<sup>34</sup> To simulate the  $C_{\text{redox}}$  of 1T- $\text{MoS}_2$ , a  $3 \times 3$  supercell was adopted, as shown in Fig. 1. The energy cutoff of plane wave was set to be 400 eV. The Monkhorst–Pack  $k$ -point grids of  $5 \times 5 \times 1$  and  $9 \times 9 \times 1$  were used for structural optimizations and self-consistent electronic structures, respectively. The optimized structure was obtained until the residual force on each atom converged to less than  $0.02 \text{ eV \AA}^{-1}$ . In the electronic



**Fig. 1** (a) Top view and (b) side view of monolayer 1T- $\text{MoS}_2$  with a layered structure. The orange, green and light green spheres represent Mo atoms, the top layer S and the bottom layer S atoms, respectively. The adsorption sites with the increase in the  $\text{H}^+$  ions coverage are labelled by numbers. (c) Top view and (d) side view of possible intrinsic vacancies: S vacancies at top layer,  $V_{S1}$ , or bottom layer,  $V_{S2}$ , and Mo vacancy,  $V_{\text{Mo}}$ .

structures steps, the energy was converged to less than  $10^{-5} \text{ eV}$ . In order to minimize the interaction between the vertically periodic layers, the thickness of the vacuum layer was set to be  $20 \text{ \AA}$  along the  $z$  direction.

## Results and discussion

### Structural configurations and stabilities of $\text{H}^+$ ions adsorbed on monolayer 1T- $\text{MoS}_2$

Monolayer 1T- $\text{MoS}_2$  exhibits a sandwiched structure, where Mo atoms are sandwiched between two layers of S atoms, as shown in Fig. 1(b). The monolayer  $\text{MoS}_2$  has two different coordinated structures, namely trigonal phase (1T) and hexagonal phase (H).<sup>35,36</sup> Compared with the H phase, the 1T- $\text{MoS}_2$  prepared by the lithium ion intercalation and hydrothermal method exhibits superior electrical conductivity and structural stability even up to  $100 \text{ }^\circ\text{C}$ .<sup>32,37,38</sup> In this study, the monolayer 1T- $\text{MoS}_2$  is our focus. The optimized lattice constant of 1T- $\text{MoS}_2$  unit cell is  $3.18 \text{ \AA}$  and the bond length of Mo–S is  $2.42 \text{ \AA}$ , which is well consistent with the experimental data and previous reports.<sup>39–41</sup>

In order to unveil the mechanism of  $C_{\text{redox}}$  in aqueous conditions, the model system of  $\text{H}^+$  ions adsorbed on the surface of monolayer 1T- $\text{MoS}_2$  was adopted. The total adsorbed system was in the neutral state since the equally additional number of electrons supplied to 1T- $\text{MoS}_2$  would compensate the  $\text{H}^+$  ions in the charging state of supercapacitors. The adsorption energy ( $E_{\text{ad}}$ ) of  $\text{H}^+$  ions on 1T- $\text{MoS}_2$  is defined as:

$$E_{\text{ad}} = (E_{\text{tot}} - E_{\text{pristine}} - nE_{\text{H}})/n \quad (1)$$

where  $E_{\text{tot}}$ ,  $E_{\text{pristine}}$ , and  $E_{\text{H}}$  are the energies of total 1T- $\text{MoS}_2$ -H supercell, pristine 1T- $\text{MoS}_2$  supercell, and one hydrogen atom, respectively, and  $n$  is the number of the adsorbed  $\text{H}^+$  ions. According to formula (1), the more negative the  $E_{\text{ad}}$ , the more stable the adsorbed system is.

For one specific  $\text{H}^+$  ion adsorbed on monolayer 1T- $\text{MoS}_2$ , three configurations of top site above S atom ( $S_{\text{top}}$ ), top site

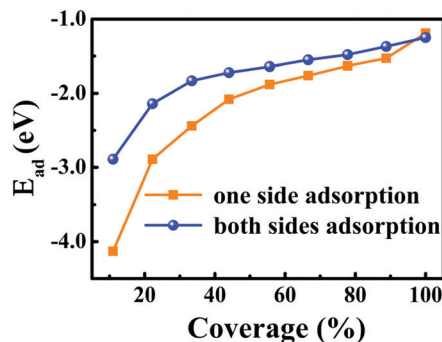


Fig. 2 Dependence of the adsorption energies of  $H^+$  ions on one side and both sides of 1T-MoS<sub>2</sub> on the  $H^+$  ions coverage. The corresponding relationship between the coverage and the number of the adsorbed  $H^+$  ions are listed in Table S1 of ESI.†

above Mo atom ( $Mo_{top}$ ) and bridge site above S-Mo bond (B) are displayed in Fig. 1(b). The adsorption energies of  $S_{top}$ ,  $Mo_{top}$  and B site are  $-4.13$  eV,  $-2.41$  eV and  $-3.75$  eV, respectively, indicating the favourable  $S_{top}$  adsorption for  $H^+$  ions. For two  $H^+$  ions adsorbed on one side of monolayer 1T-MoS<sub>2</sub>, two possible configurations of the nearest neighbouring (NN) and next nearest neighbouring adsorptions existed based on the distance between these two  $H^+$  ions. Through comparing the  $E_{ad}$  of the NN and NNN configurations, we found that the NNN adsorption was more stable with an energy difference of 0.15 eV. With the increase in the adsorbed  $H^+$  ions, the possible configurations and the  $E_{ad}$  were also examined, as shown in Table S1 of the ESI.† The most stable adsorbed sites are labelled by numbers in Fig. 1(a) according to the coverage

increase in the  $H^+$  ions. It can be seen that the  $H^+$  ions tend to distribute away from each other to minimize the coulomb repulsions among them when the coverage is not large. To provide references for experiments, the coverage is used and defined as the ratio of the adsorbed  $H^+$  ions amount to the total active  $S_{top}$  sites. As the coverage increases, the absolute values of  $E_{ad}$  gradually decrease but always with negative values in Fig. 2, implying the stable adsorbed state up to 100% coverage.

Besides the one side adsorption of  $H^+$  ions on 1T-MoS<sub>2</sub>, both sides adsorption is reasonable to improve the  $C_{redox}$  in practical nano array electrodes.<sup>42–46</sup> Through comparing the  $E_{ad}$  for different adsorbed configurations, it is seen that the symmetrically adsorbed configurations between the top and the bottom S layers are favourable. As displayed in Fig. 2, the  $E_{ad}$  of  $H^+$  ions on both sides also follows the similar character except for slightly weaker than that of the one side adsorption, which is basically induced by the stronger coulomb repulsions among more adsorbed  $H^+$  ions.

### Electronic structures of the $H^+$ ions adsorption on monolayer 1T-MoS<sub>2</sub>

For supercapacitor electrodes, the fast charge/discharge rate is closely related to the carrier conductivity, which is reflected by the electronic structures of the electrode materials. For one side  $H^+$  adsorption of monolayer 1T-MoS<sub>2</sub>, the density of states (DOS) and partial DOS (PDOS) of monolayer 1T-MoS<sub>2</sub> under the different adsorption coverages of  $H^+$  ions are presented in Fig. 3. The evolution of band structures with the  $H^+$  coverage is also provided in Fig. S1 of ESI.† In Fig. 3(a), the pristine monolayer 1T-MoS<sub>2</sub> shows metallic character, which is consistent with the previous report.<sup>39</sup> The states near the Fermi level ( $E_F$ ) are

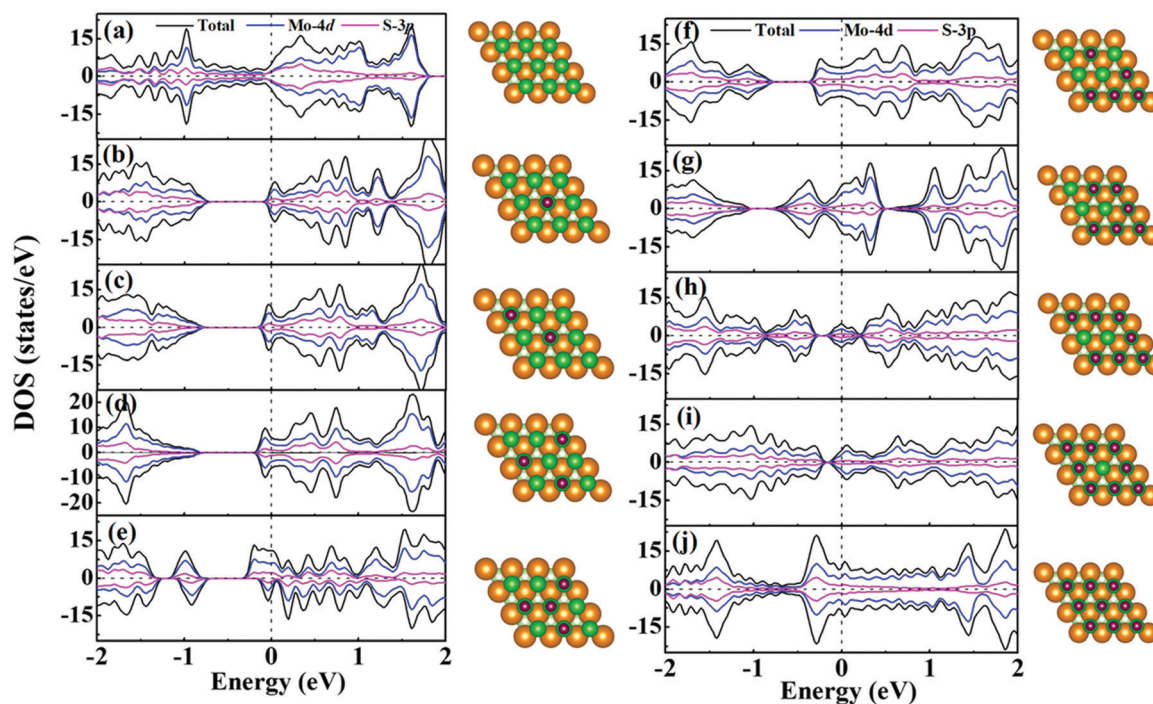


Fig. 3 Total DOS, PDOS and geometric structures of 1T-MoS<sub>2</sub> under different adsorption coverage of  $H^+$  ions on one side of 1T-MoS<sub>2</sub>. Pristine 1T-MoS<sub>2</sub> in (a), and the coverage is 11.1% in (b), 22.2% in (c), 33.3% in (d), 44.4% in (e) 55.6% in (f) 66.7% in (g) 77.8% in (h), 88.9% in (i), 100% in (j). The Fermi level is set to be 0.0 eV.

basically contributed by Mo-4d and S-3p orbitals due to the p-d hybridizations. When the active  $S_{\text{top}}$  sites on one side of monolayer 1T-MoS<sub>2</sub> partially adsorb the H<sup>+</sup> ions, the monolayer 1T-MoS<sub>2</sub> opens one band gap and the band gap value decreases with the increment in the H<sup>+</sup> coverage in Fig. 3(b-i). Under these situations, the  $E_{\text{F}}$  is located within the conduction band, and the monolayer 1T-MoS<sub>2</sub> still displays the metallic-like character. It is noted that the monolayer 1T-MoS<sub>2</sub> demonstrates excellent metallic properties under the full coverage of H<sup>+</sup> ions in Fig. 3(j) and Fig. S1(i) (ESI<sup>†</sup>). For monolayer 1T-MoS<sub>2</sub> with H<sup>+</sup> ions adsorbed on both sides, the DOS, PDOS and band structures of the 1T-MoS<sub>2</sub> under different adsorption coverages are shown in Fig. S2 and S3 (ESI<sup>†</sup>). It is found that the monolayer 1T-MoS<sub>2</sub> with both sides adsorption also shares the similar character to that with one side adsorption. Particularly, the metallic band structures of monolayer 1T-MoS<sub>2</sub> are well preserved at 100% coverage of H<sup>+</sup> ions in Fig. S3(i) (ESI<sup>†</sup>). Combined with the practical applications, the larger adsorption coverage of H<sup>+</sup> ions and the excellent electrical conductivity are desirable to induce the promising large  $C_{\text{redox}}$  for monolayer 1T-MoS<sub>2</sub>.

Since the  $C_{\text{redox}}$  is associated with the charge transfer between the adsorbed ions and the electrode, the charge density difference  $\Delta\rho$  is presented in Fig. 4 to more vividly observe the charge transfer between the H<sup>+</sup> ions and the charging monolayer 1T-MoS<sub>2</sub>.  $\Delta\rho$  is expressed as

$$\Delta\rho = \rho(\text{MoS}_2\text{-}n\text{H}) - \rho(\text{MoS}_2^{n-}) - \rho(n\text{H}^+) \quad (2)$$

where  $\rho(\text{MoS}_2\text{-}n\text{H})$ ,  $\rho(\text{MoS}_2^{n-})$  and  $\rho(n\text{H}^+)$  are the electron densities of 1T-MoS<sub>2</sub>- $n\text{H}$  systems, the charging monolayer 1T-MoS<sub>2</sub> and the isolated H<sup>+</sup> ions.

As shown in Fig. 4, the monolayer 1T-MoS<sub>2</sub> donates electrons to H<sup>+</sup> ions in all coverage cases. In detail, the localized charge transfer basically occurs between the H<sup>+</sup> ions and adjacent Mo atoms in the lower coverage systems. With the coverage increase in the H<sup>+</sup> ions, the charge density difference distribution tends

to be delocalized within 2D sheet of 1T-MoS<sub>2</sub>, indicating the improvement in the metallic properties for monolayer 1T-MoS<sub>2</sub>. In particular, the charge density difference in Fig. 4(i) is almost uniformly distributed in the 2D plane of the monolayer 1T-MoS<sub>2</sub>, consistent with the metallic band structures of 1T-MoS<sub>2</sub> with 100% coverage of H<sup>+</sup> ions in Fig. 3(j).

### Pseudocapacitance of H<sup>+</sup> ions adsorption on monolayer 1T-MoS<sub>2</sub>

Due to the fast reversible adsorption/desorption process of H<sup>+</sup> ions in supercapacitors, the electrode of monolayer 1T-MoS<sub>2</sub> would gain electrons to saturate the positive charges of H<sup>+</sup> ions in the charging state. The  $C_{\text{redox}}$  induced in this process can be calculated by

$$C_{\text{redox}} = \Delta Q/\Delta V \quad (3)$$

where  $\Delta Q$  and  $\Delta V$  are the charge transfer from the electrode to H<sup>+</sup> ions and the potential change of the electrode, respectively.  $\Delta Q$  can be obtained based on the Bader charge analysis and the potential change in monolayer 1T-MoS<sub>2</sub> is estimated by work function (WF) change:

$$\Delta V = \Delta\text{WF}/e \quad (4)$$

$$\Delta\text{WF} = \text{WF}_{\text{MoS}_2\text{-}n\text{H}} - \text{WF}_{\text{MoS}_2^{\Delta Q^-}} \quad (5)$$

where  $\text{WF}_{\text{MoS}_2\text{-}n\text{H}}$  and  $\text{WF}_{\text{MoS}_2^{\Delta Q^-}}$  are the work functions of the monolayer 1T-MoS<sub>2</sub> adsorbed H<sup>+</sup> ions at neutral state and the charged monolayer 1T-MoS<sub>2</sub> with negative charge of  $\Delta Q$ , respectively. For monolayer 1T-MoS<sub>2</sub> itself, its  $E_{\text{F}}$  crosses the energy bands and it is metallic whether in the pristine state or in charging state with additional negative charges through external circuit, as demonstrated by the uniformly distributed charges on the 2D plane of 1T-MoS<sub>2</sub> in Fig. S4 of ESI<sup>†</sup>. Thus, the WF of monolayer 1T-MoS<sub>2</sub> in charging state is accessible *via* the shift of  $E_{\text{F}}$  ( $\Delta E_{\text{F}}$ ) relative to pristine 1T-MoS<sub>2</sub> according to the rigid-band approximation, which is presented in Fig. S5 of ESI<sup>†</sup>.

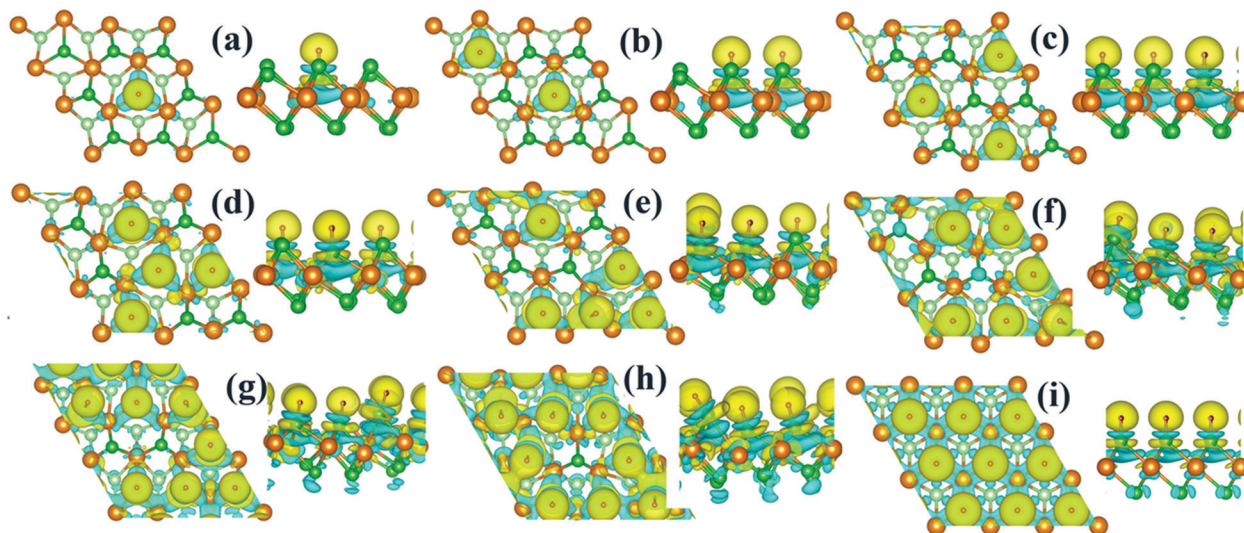


Fig. 4 Top and side views of charge density difference distribution. The coverage is 11.1% in (a), 22.2% in (b), 33.3% in (c), 44.4% in (d) 55.6% in (e) 66.7% in (f) 77.8% in (g), 88.9% in (h), and 100% in (i). The yellow and cyan isosurfaces with an isovalue of  $0.003 \text{ e} \text{ \AA}^{-3}$  correspond to charge accumulation and depletion, respectively.

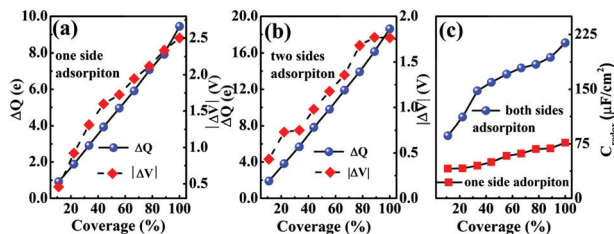


Fig. 5 Charge transfer and potential change at different  $H^+$  ion coverages on one side (a) and on both sides (b) of monolayer 1T-MoS<sub>2</sub>. Pseudocapacitance corresponding to (a) and (b) is presented in (c).

In Fig. 5(a and b), it is seen that the charge transfer,  $\Delta Q$ , from the charged electrode to the  $H^+$  ions increases almost linearly with the increase in the coverage of  $H^+$  ions on one side or both sides of monolayer 1T-MoS<sub>2</sub>. At the same time, the WF values of monolayer 1T-MoS<sub>2</sub> adsorbed by  $H^+$  ions at neutral state,  $WF_{MoS_2-nH}$ , and monolayer 1T-MoS<sub>2</sub> with negatively charges,  $WF_{MoS_2\Delta Q^-}$ , gradually decrease, which are listed in Table S2 of ESI.† The former is associated with the charge transfer and the surface electrical dipole,<sup>47</sup> and the latter is related to  $\Delta E_F$  of the charged 1T-MoS<sub>2</sub> relative to pristine 1T-MoS<sub>2</sub>. Due to the extended DOS near the  $E_F$  of monolayer 1T-MoS<sub>2</sub>,  $\Delta E_F$  slightly increases with the increase in the electrons gain, leading to a slow reduction in  $WF_{MoS_2\Delta Q^-}$ . Hence, the potential change is related to the work function variation of monolayer 1T-MoS<sub>2</sub> after adsorption by  $H^+$  ions. As the rate of decrease in  $WF_{MoS_2-nH}$  gets slower with the increase in the  $H^+$  ions coverage, particularly close to the complete coverage, the maximum values of the  $C_{redox}$ ,  $\sim 76.67 \mu F cm^{-2}$  ( $\sim 252.81 F g^{-1}$ ) for one side adsorption and  $\sim 213.67 \mu F cm^{-2}$  ( $704.53 F g^{-1}$ ) for two sides adsorption, respectively, have been obtained in Fig. 5(c). Our theoretically obtained  $C_{redox}$  is higher than the reported experimental data, and further improvement of  $C_{redox}$  would be expected and promising in experiments.<sup>29,48,49</sup>

### Pseudocapacitance manipulation of monolayer 1T-MoS<sub>2</sub> by intrinsic defects

The maximum  $C_{redox}$  of monolayer 1T-MoS<sub>2</sub> is achieved at full coverage of  $H^+$  ions due to the well-preserved structures and metallic electronic structures of monolayer 1T-MoS<sub>2</sub>. Combined with the experiments, the intrinsic defects of S vacancies or Mo vacancies possibly exist in the growth process due to different preparation conditions.<sup>50–53</sup> To illustrate the influence of the defects on the  $C_{redox}$  and to further manipulate  $C_{redox}$  through introducing defects conversely, the effects of possible vacancies are explored in 100% coverage of  $H^+$  ions on monolayer 1T-MoS<sub>2</sub>. As shown in Fig. 1(c and d), for the  $3 \times 3$  supercell with 100% coverage of  $H^+$  ions on one side and on both sides, one Mo vacancy ( $V_{Mo}$ ), one S vacancy ( $V_{S1}$ ), and two S vacancies in top and bottom layers ( $V_{S1} + V_{S2}$ ) are considered. Under this situation, the  $E_{ad}$  of  $H^+$  ions is expressed as,

$$E_{ad} = (E_{tot} - E_{defect} - nE_H)/n \quad (6)$$

where  $E_{tot}$ ,  $E_{defect}$ , and  $E_H$  are the energies of the defective 1T-MoS<sub>2</sub>- $nH$  system, the defective 1T-MoS<sub>2</sub> system, and one

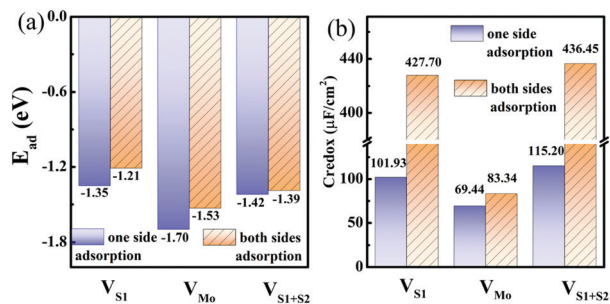


Fig. 6 (a) Adsorption energies and (b) pseudocapacitance of monolayer 1T-MoS<sub>2</sub> with 100% coverage of  $H^+$  ions on one side and both sides under different intrinsic defect states.

hydrogen atom, respectively. Moreover,  $n = 9$  or  $n = 18$  corresponds to the 100% coverage of  $H^+$  ions on one side or both sides of monolayer 1T-MoS<sub>2</sub>.

With the intrinsic defects in 1T MoS<sub>2</sub>- $nH$  system, the  $E_{ad}$  of  $H^+$  ions is still negative for one side and both sides adsorption cases in Fig. 6(a), implying the energetically stable adsorption of  $H^+$  ions on defect monolayer 1T-MoS<sub>2</sub>. Moreover, the adsorption of  $H^+$  ions on 1T-MoS<sub>2</sub> in the presence of  $V_{S1}$ ,  $V_{Mo}$ , or  $V_{S1+S2}$  gets slightly stronger than that on pristine 1T-MoS<sub>2</sub>, which is resulted from the nearest  $H^+$  moving closer to the vacancy site and the enhanced interaction between the  $H^+$  and the 1T-MoS<sub>2</sub>. In Fig. 6(a), it is also found that the adsorption of  $H^+$  ions on one side is more stable than that on both sides owing to the stronger Coulomb repulsion interaction between more  $H^+$  ions adsorbed on both sides of monolayer 1T-MoS<sub>2</sub>. In terms of the DOS in Fig. S6 of ESI,† all the above-mentioned systems demonstrate metallic characters, which could ensure the defective monolayer 1T-MoS<sub>2</sub> possesses the well rate performance as an electrode.

In the presence of S vacancies, the  $C_{redox}$  is increased to  $101.93 \mu F cm^{-2}$  ( $336.09 F g^{-1}$ ) and  $115.20 \mu F cm^{-2}$  ( $379.83 F g^{-1}$ ) in 100% coverage of  $H^+$  ions on one side of monolayer 1T-MoS<sub>2</sub> with  $V_{S1}$  and  $V_{S1+S2}$ , respectively, as shown in Fig. 6(b). Furthermore, the  $C_{redox}$  in 100% coverage of  $H^+$  ions on both sides of monolayer 1T-MoS<sub>2</sub> system is enhanced to  $427.70 \mu F cm^{-2}$  ( $1410.24 F g^{-1}$ ) and  $436.45 \mu F cm^{-2}$  ( $1439.08 F g^{-1}$ ) when  $V_{S1}$  and  $V_{S1+S2}$  are included. It is observed from Table S4 (ESI†) that the  $WF_{MoS_2-nH}$  of the system with S vacancies with respect to that of the intrinsic system is increased, which is induced by more electrons transfer from the defect 1T-MoS<sub>2</sub> to  $H^+$  ions. At the same time,  $WF_{MoS_2\Delta Q^-}$  of monolayer 1T-MoS<sub>2</sub> with S vacancies at charging state is suppressed relative to that of the pristine 1T-MoS<sub>2</sub> since the S vacancies contribute more electrons to facilitate the electrons emission. Thus, the potential difference,  $\Delta V$ , is also suppressed, leading to an increased  $C_{redox}$  in Fig. 6(b).

In contrast, the presence of the  $V_{Mo}$  reduces the  $C_{redox}$  of monolayer 1T-MoS<sub>2</sub>. For instance, the  $C_{redox}$  of the systems with one-side adsorption and two-side adsorption of  $H^+$  ions are  $69.44 \mu F cm^{-2}$  and  $83.34 \mu F cm^{-2}$ , respectively, while the corresponding values are  $76.67 \mu F cm^{-2}$  and  $213.67 \mu F cm^{-2}$  in pristine systems with the same  $H^+$  ion coverage. As indicated

in Table S4 (ESI<sup>†</sup>), the  $V_{\text{Mo}}$  lifts the WF of pristine monolayer 1T-MoS<sub>2</sub> ( $WF_{\text{MoS}_2^{\Delta Q^-}}$ ) and simultaneously suppresses the WF of the total system ( $WF_{\text{MoS}_2-n\text{H}}$ ), generating a larger potential difference between them and a decreased  $C_{\text{redox}}$ . Through looking into the effects of the intrinsic defects on the  $C_{\text{redox}}$  of monolayer 1T-MoS<sub>2</sub>, the  $C_{\text{redox}}$  would be manipulated and improved to a large extent by controlling the experimental growth conditions in S poor environment to produce S vacancies.

In addition, about the approach to calculate the  $C_{\text{redox}}$ ,  $C_{\text{redox}}$  is directly dependent on  $\Delta Q$  and  $\Delta V$ , which can be applied in other 2D materials and oxides.  $\Delta Q$  is acquired by the charge transfer based on Bader charge analysis, and  $\Delta V$  is obtained by the work function change of the electrode material. In other electrodes materials, the stability of the adsorbed systems, the charge transfer between the electrodes and the adsorbed ions, and the work function change are possibly different from monolayer 1T-MoS<sub>2</sub> system. Particularly, the optimal performance of  $C_{\text{redox}}$  and the favourable coverage of H<sup>+</sup> ions need to be further examined.

## Conclusions

In summary, the correlation between the  $C_{\text{redox}}$  and the microscopic electronic structures was examined based on the DFT calculations. One factor determining the  $C_{\text{redox}}$  is the charge transfer from the electrode to the adsorbed H<sup>+</sup> ions. The other is the potential change of the electrode due to the H<sup>+</sup> ions adsorption. For pristine monolayer 1T-MoS<sub>2</sub>, the amount of the charge transfer increases almost linearly with the increase in the H<sup>+</sup> ion coverage. Thus, the  $C_{\text{redox}}$  is mainly dependent on the potential change of the monolayer 1T-MoS<sub>2</sub>. The smaller decrease rate of the potential change in monolayer 1T-MoS<sub>2</sub> with fully coverage of H<sup>+</sup> ions on both sides results in a higher  $C_{\text{redox}}$  of 213  $\mu\text{F cm}^{-2}$  (704  $\text{F g}^{-1}$ ). Moreover, the  $C_{\text{redox}}$  is much improved after introducing S vacancies. These results would provide theoretical guidance in understanding and manipulating the pseudocapacitance properties of 1T-MoS<sub>2</sub> or other similar quasi-two-dimensional materials for further experimental explorations.

## Author contributions

Wei-Hua Wang and Weichao Wang proposed the research idea, wrote and reviewed the manuscript. Zhenzhou Zhang performed the calculations and wrote the draft. Maokun Wu, Lijing Wang and Jin Wang did part of calculations. All the authors devoted to the discussions and revised the manuscript.

## Conflicts of interest

There are no conflicts to declare.

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## References

- 1 L. F. Chen, Y. Feng, H. W. Liang, Z. Y. Wu and S. H. Yu, *Adv. Energy Mater.*, 2017, 7, 1700826.
- 2 C. P. Grey and J. M. Tarascon, *Nat. Mater.*, 2017, 16, 45–56.
- 3 G. Zhang, Y. Han, C. Shao, N. Chen, G. Sun, X. Jin, J. Gao, B. Ji, H. Yang and L. Qu, *Mater. Chem. Front.*, 2018, 2, 1750.
- 4 N. Choudhary, M. A. Islam, J. H. Kim, T. J. Ko, A. Schropp, L. Hurtado, D. Weitzman, L. Zhai and Y. Jung, *Nano Today*, 2018, 19, 16–40.
- 5 N. Yang, Y. Zhang, J. E. Halpert, J. Zhai, D. Wang and L. Jiang, *Small*, 2012, 8, 1762–1770.
- 6 Y. Zhao, H. Tang, N. Yang and D. Wang, *Adv. Sci.*, 2018, 5, 1800959.
- 7 F. Wang, X. Wu, X. Yuan, Z. Liu, Y. Zhang, L. Fu, Y. Zhu, Q. Zhou, Y. Wu and W. Huang, *Chem. Soc. Rev.*, 2017, 46, 6816–6854.
- 8 R. Li, Y. Wang, C. Zhou, C. Wang, X. Ba, Y. Li, X. Huang and J. Liu, *Adv. Funct. Mater.*, 2015, 25, 5384–5394.
- 9 F. Yu, T. Wang, Z. Wen and H. Wang, *J. Power Sources*, 2017, 364, 9–15.
- 10 H. Wang, C. M. B. Holt, Z. Li, X. Tan, B. S. Amirkhiz, Z. Xu, B. C. Olsen, T. Stephenson and D. Mitlin, *Nano Res.*, 2012, 5, 605–617.
- 11 C. Liu, Z. Yu, D. Neff, A. Zhamua and B. Z. Jang, *Nano Lett.*, 2010, 10, 4863–4868.
- 12 B. E. Conway, *J. Electrochem. Soc.*, 1991, 138, 1539.
- 13 K. S. Novoselov, A. K. Geim, S. V. Morozov, D. Jiang, Y. Zhang, S. V. Dubonos, I. V. Grigorieva and A. A. Firsov, *Science*, 2004, 306, 666–669.
- 14 N. Yang, J. Zhai, D. Wang, Y. Chen and L. Jiang, *ACS Nano*, 2010, 2, 887–894.
- 15 H. Ren, H. Shao, L. Zhang, D. Guo, Q. Jin, R. Yu, L. Wang, Y. Li, Y. Wang, H. Zhao and D. Wang, *Adv. Energy Mater.*, 2015, 5, 1500296.
- 16 Y. Zhao, J. Wan, H. Yao, L. Zhang, K. Lin, L. Wang, N. Yang, D. Liu, L. Song, J. Zhu, L. Gu, L. Liu, H. Zhao, Y. Li and D. Wang, *Nat. Chem.*, 2018, 10, 924–931.
- 17 N. Yang, Y. Liu, H. Wen, Z. Tang, H. Zhao, Y. Li and D. Wang, *ACS Nano*, 2013, 7, 1504–1512.
- 18 J. Yan, C. E. Ren, K. Maleski, C. B. Hatter, B. Anasori, P. Urbankowski, A. Sarycheva and Y. Gogotsi, *Adv. Funct. Mater.*, 2017, 27, 1701264.
- 19 H. Wang, T. Maiyalagan and X. Wang, *ACS Catal.*, 2012, 2, 781–794.
- 20 T. Akhter, M. M. Islam, S. N. Faisal, E. Haque, A. I. Minett, H. K. Liu, K. Konstantinov and S. X. Dou, *ACS Appl. Mater. Interfaces*, 2016, 8, 2078–2087.
- 21 T. Lin, I.-W. Chen, F. Liu, C. Yang, H. Bi, F. Xu and F. Huang, *Science*, 2015, 350, 6267.
- 22 Y. Xu, C. Y. Chen, Z. Zhao, Z. Lin, C. Lee, X. Xu, C. Wang, Y. Huang, M. I. Shakir and X. Duan, *Nano Lett.*, 2015, 15, 4605–4610.

- 23 Z. Teng, N. Yang, H. Lv, S. Wang, M. Hu, C. Wang, D. Wang and G. Wang, *Chem*, 2019, **5**, 1–17.
- 24 N. Yang and D. Wang, *Chem*, 2018, **4**, 2260–2277.
- 25 Q. Tang, Z. Zhou and P. Shen, *J. Am. Chem. Soc.*, 2012, **134**, 16909–16916.
- 26 M. Naguib, M. Kurtoglu, V. Presser, J. Lu, J. J. Niu, M. Heon, L. Hultman, Y. Gogotsi and M. W. Barsoum, *Adv. Mater.*, 2011, **23**, 4248–4253.
- 27 H. W. Wang, M. Naguib, K. Page, D. J. Wesolowski and Y. Gogotsi, *Chem. Mater.*, 2016, **28**, 349–359.
- 28 B. Anasori, M. R. Lukatskaya and Y. Gogotsi, *Nat. Rev. Mater.*, 2017, **2**, 16098.
- 29 X. Geng, Y. Zhang, Y. Han, J. Li, L. Yang, M. Benamara, L. Chen and H. Zhu, *Nano Lett.*, 2017, **17**, 1825–1832.
- 30 M. Acerce, D. Voiry and M. Chhowalla, *Nat. Nanotechnol.*, 2015, **10**, 313–317.
- 31 X. Cong, C. Cheng, Y. Liao, Y. Ye, C. Dong, H. Sun, X. Ji, W. Zhang, P. Fang, L. Miao and J. Jiang, *J. Phys. Chem. C*, 2015, **119**, 20864–20870.
- 32 Q. Tang and D. Jiang, *Chem. Mater.*, 2015, **27**, 3743–3748.
- 33 G. Kresse and J. Furthmüller, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1996, **54**, 11169–11186.
- 34 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865.
- 35 M. Calandra, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2013, **88**, 245428.
- 36 J. Heising and M. G. Kanatzidis, *J. Am. Chem. Soc.*, 1999, **121**, 638–643.
- 37 H. L. Tsai, J. Heising, J. L. Schindler, C. R. Kannewurf and M. G. Kanatzidis, *Chem. Mater.*, 1997, **9**, 879–882.
- 38 F. Wypych and R. Schollhorn, *J. Chem. Soc., Chem. Commun.*, 1992, **19**, 1386–1388.
- 39 T. Hu, R. Li and J. Dong, *J. Chem. Phys.*, 2013, **139**, 174702.
- 40 M. Kan, J. Y. Wang, X. W. Li, S. H. Zhang, Y. W. Li, Y. Kawazoe, Q. Sun and P. Jena, *J. Phys. Chem. C*, 2014, **118**, 1515–1522.
- 41 K. E. Dungey, M. D. Curtis and J. E. P. Hahn, *Chem. Mater.*, 1998, **10**, 2152–2161.
- 42 X. Yang, L. Zhao and J. Lian, *J. Power Sources*, 2017, **343**, 373–382.
- 43 M. Wang, L. Fan, D. Tian, X. Wu, Y. Qiu, C. Zhao, B. Guan, Y. Wang, N. Zhang and K. Sun, *ACS Energy Lett.*, 2018, **3**, 1627–1633.
- 44 L. Wang, X. Zhang, Y. Ma, M. Yang and Y. Qi, *J. Phys. Chem. C*, 2017, **121**, 9089–9095.
- 45 J. Wang, D. Chao, J. Liu, L. Li, L. Lai, J. Lin and Z. Shen, *Nano Energy*, 2014, **7**, 151–160.
- 46 W. J. Tang, X. L. Wang, D. Xie, X. H. Xia, C. D. Gu and J. P. Tu, *J. Mater. Chem. A*, 2018, **6**, 18318–18324.
- 47 M. Khazaei, M. Arai, T. Sasaki, A. Ranjbar, Y. Liang and S. Yunoki, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2015, **92**, 075411.
- 48 N. Joseph, P. M. Shafi and A. C. Bose, *New J. Chem.*, 2018, **42**, 12082–12090.
- 49 C. Y. Zhao, J. M. Ang, Z. L. Liu and X. Lu, *Chem. Eng. J.*, 2017, **330**, 462–469.
- 50 R. Addou, L. Colombo and R. M. Wallace, *ACS Appl. Mater. Interfaces*, 2015, **7**, 11921–11929.
- 51 Z. Q. Fan, X. W. Jiang, J. W. Luo, L. Y. Jiao, R. Huang, S. S. Li and L. W. Wang, *Phys. Rev. B*, 2017, **96**, 165402.
- 52 W. Zhou, X. L. Zou, S. Najmaei, Z. Liu, Y. Shi, J. Kong, J. Lou, P. M. Ajayan, B. I. Yakobson and J. C. Idrobo, *Nano Lett.*, 2013, **13**, 2615–2622.
- 53 J. Hong, Z. Hu, M. Probert, K. Li, D. Lv, X. Yang, L. Gu, N. Mao, Q. Feng, L. Xie, J. Zhang, D. Wu, Z. Zhang, C. Jin, W. Ji, X. Zhang, J. Yuan and Z. Zhang, *Nat. Commun.*, 2015, **6**, 6293.