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MINIREVIEW

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Oxide-based nanomaterials for fuel cell catalysis: the interplay between supported single Pt atoms and particles

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The concept of single atom catalysis offers maximum noble metal efficiency for the development of low-cost catalytic materials. Among possible applications are catalytic materials for proton exchange membrane fuel cells. In the present review, recent efforts towards the fabrication of single atom catalysts on nanostructured ceria and their reactivity are discussed in the prospect of their employment as anode catalysts. The remarkable performance and the durability of the ceria-based anode catalysts with ultra-low Pt loading result from the interplay between two states associated with supported atomically dispersed Pt and sub-nanometer Pt particles. The occurrence of these two states is a consequence of strong interactions between Pt and nanostructured ceria that yield atomically dispersed species under oxidizing conditions and sub-nanometer Pt particles under reducing conditions. The square-planar arrangement of four O atoms on {100} nanoterraces has been identified as the key structural element on the surface of the nanostructured ceria where Pt is anchored in the form of Pt²⁺ species. The conversion of Pt²⁺ species to sub-nanometer Pt particles is triggered by a redox process involving Ce³⁺ centers. The latter emerge due to either oxygen vacancies or adsorption of reducing agents. The unique properties of the sub-nanometer Pt particles arise from metal-support interactions involving charge transfer, structural flexibility, and spillover phenomena. The abundance of specific adsorption sites similar to those on {100} nanoterraces determines the ideal (maximum) Pt loading in Pt-CeO_x films that still allows reversible switching between the atomically dispersed Pt and sub-nanometer particles yielding high activity and durability during fuel cell operation.

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1. Introduction

Hydrogen powered proton exchange membrane fuel cells (PEMFCs) are considered potential next generation power

sources for a variety of small to medium-scale applications, including automotive vehicle propulsion and chip-integrated micro-devices.^{1–3} For PEMFC technology, platinum is the essential catalytic element. Typically, commercial anode catalysts contain about 2–5 mg cm⁻² of the noble metal. The high cost of platinum is the main factor limiting the large-scale application of fuel cell technology. Therefore, great efforts are dedicated to the development of catalytic materials for PEMFCs to meet the standards defined by the US Department of Energy (DOE).⁴ The latest developments in fuel cell technology may reduce the total Pt loading in PEMFCs to 0.15 mg cm⁻² which is still above the target value (0.125 mg cm⁻²) set for the year 2017.⁴

In the prospect of further reduction of Pt loading, supported single atom catalysts (SACs) offer ultimate noble metal efficiency by exposing the entire noble-metal content to reactants.^{5–8} The synthesis and performance of SACs in numerous heterogeneous reactions have been recently reviewed in detail.^{6–10} Several key challenges have been identified with respect to the stability and reactivity of the atomically dispersed noble metals. In particular, anchoring of noble metal atoms at appropriate supports requires the presence of

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specific adsorption sites that are capable of stabilizing metal atoms against sintering and agglomeration into particles. Such sites have been identified on the surfaces of graphene,^{11–13} nitrides,^{14–18} zeolites,¹⁹ and metal oxides.^{20–31} In most cases, atomically dispersed metals are found to be anchored in surface cationic positions of the host oxide.^{23,28,30,32} Depending on the oxidation state of the anchored metal atom, spontaneous formation of anionic or cationic vacancies occurs to balance the charge. In some cases, anchoring of the metal atom is achieved through the spatial confinement in pores or open channels of the support.^{17,18,33} The stability of the anchored metal atoms and their propensity to agglomerate is determined by the energy difference between the states associated with an adsorbed metal atom and a supported metal particle. In this respect, the use of metal oxides allows employing metal–support interactions to achieve a strong binding between the metal and the support.^{34–36} Conceptually, oxide supported noble metal SACs are closely related to noble metal-doped oxides.^{36–38} However, the common disadvantage of the latter is that a substantial amount of the noble metal is atomically dispersed in the bulk.^{36–38} As a result, the density of the noble-metal sites at the surface is low. An increase of the noble metal loading in doped oxides, on the other hand, often gives rise to metallic particles at the surface.³⁶

Different strategies are applied to increase the density of the atomically dispersed species at the surface. These include grafting of the atomically dispersed metals, the use of stabilizing ligands, and nanostructuring of the support. In particular, new sites emerge at the surface of the nanostructured oxides that may anchor atomically dispersed noble metals in a structural environment which is energetically highly favorable²¹ with respect to cationic substitution in the doped bulk.³⁶ Under these circumstances, the segregation of the noble metal is driven thermodynamically, yielding a high density of atomically dispersed noble metal at the surface.²⁷

In the following review we summarize the properties of new materials containing atomically dispersed noble metals anchored at surface sites of nanostructured ceria. In particu-

lar, the catalytic performance of atomically dispersed platinum supported on nanostructured ceria is discussed in the prospect of applications as anode catalysts for PEMFCs. The remarkable properties of these catalysts involve the interplay between the two states associated with the atomically dispersed noble metal and sub-nanometer particles. The great stability and durability of the prototype anode catalyst arise from reversible switching between these two states under operating PEMFC conditions.

2. Supported single atom catalysts in proton exchange membrane fuel cells

2.1. High efficiency at ultra-low noble metal loading

The performance of Pt–CeO₂ catalysts with ultra-low Pt loading prepared by means of radio frequency sputtering was tested under relevant PEMFC conditions.^{21,39–44} The key parameter determining the cost efficiency of the fuel cell catalyst is the specific power (SP), *i.e.* the power density (PD) per weight of noble metal. The corresponding SPs and PDs achieved using the Pt–CeO₂ catalysts at the anode (Table 1) were compared with those of a commercial Pt nanoparticle catalyst and Pt thin films (Table 2) under identical operation conditions.

The thin film of the Pt–CeO₂ catalyst deposited directly on the gas diffusion layer (GDL) yielded a higher PD with respect to the reference PtRu anode catalyst despite the ten-fold lower Pt loading.^{39,40} The use of double-wall (DWCNT),⁴⁰ multi-wall (MWCNT),⁴² and chemical vapor deposited (CVD-CNT)⁴³ carbon nanotubes resulted in a further increase of PDs and SPs (see Table 1). With respect to the commercial Pt anode (Table 2a), the Pt–CeO₂ catalyst yielded an around 10²-fold increase of the SP.⁴¹

Additionally, an increase of the operation temperature of the fuel cell yielded an approximately two-fold increase of the PD and SP.⁴³ However, the highest PD and SP were obtained with the Pt–CeO₂ catalyst deposited on carbon nanoparticle coated GDL (nanoGDL, also n-GDL).^{21,41,44}

Table 1 Power density (PD) and specific power (SP) obtained with the Pt–CeO₂ anode catalyst as a function of Pt loading, the catalyst support, and temperature

Catalyst support	Pt loading ($\mu\text{g cm}^{-2}$)	T (K)	PD (mW cm^{-2})	SP (W mg^{-1})	Ref.
GDL	2.2	300	4.9	2.5	39
	2.2 ^a	300	10.5	5.4	39
	1.2	300	12.3	10	40
	1.2	313	15.5	12.6	40
	0	300	0.41	—	40
DWCNT/GDL	1.2	300	43	35	40
MWCNT/GDL	0.9	300	25	28	42
	0.9	348	38	42	42
CVD-CNT/GDL	0.9	300	40	44.4	43
		342	70	77.8	43
		348	74	82.2	43
n-GDL	4	338	330	82.5	41, 44
	2	338	410	205	21, 41
	0.6	338	170	283	41

^a The catalyst contains Sn.



Table 2 Power density (PD) and specific power (SP) obtained with anode catalysts based on a commercial Pt nanoparticle catalyst (a) and thin Pt films (b and c) as a function of noble metal loading, the catalyst support, and temperature

Anode catalyst	Pt loading ($\mu\text{g cm}^{-2}$)	T (K)	PD (mW cm^{-2})	SP (W mg^{-1})	Ref.	
(a)	Pt/GDL	2000	338	440	0.22	41
(b)	Pt/GDL	21.45	300	2.4	0.12	40
(c)	Pt/n-GDL	2	338	75	37	41

The morphology of the GDL support has a critical influence on the performance of the Pt–CeO₂ catalyst. Scanning electron microscopy (SEM) revealed characteristic differences in the structure of the bare GDL, GDL coated with DWCNTs, MWCNTs, and CVD-CNTs, and n-GDL (see Fig. 1).

The main difference between the CNT-coated GDL supports (Fig. 1b–d) is the diameter of the tubes and their alignment with respect to the surface of GDL. In particular, CVD-CNTs grow perpendicular to the surface of GDL while DWCNTs and MWCNTs deposited by spin-coating are aligned parallel to the surface of GDL.⁴²

2.2. The composition of the catalysts

The composition of the Pt–CeO₂ anode catalysts and the Pt oxidation state were investigated as a function of the information depth.^{39,40} The experimental approach involved a combination of high-resolution synchrotron spectroscopy (SRPES), angle-resolved X-ray photoelectron spectroscopy (AR XPS), and hard X-ray photoelectron spectroscopy (HAXPS).⁴⁸ The information depth achieved with these techniques increases with increasing photon energies ($h\nu$) and $\cos(\alpha)$ with respect to the surface normal from 0.5 nm (SRPES) and 1–2 nm (AR XPS) to 7 nm (HAXPS). The corresponding Pt 4f spectra obtained from the Pt–CeO₂ thin films with SRPES, AR XPS, and HAXPS are shown in Fig. 2.

The major contributions in the Pt 4f spectra arise from atomically dispersed Pt²⁺ and Pt⁴⁺ species. Additionally, traces of metallic Pt⁰ were identified at the surface of Pt–CeO₂ films (Fig. 2, bottom spectrum). The formation of ionic species is typically found for films prepared by magnetron co-sputtering of CeO₂ in combination with transition

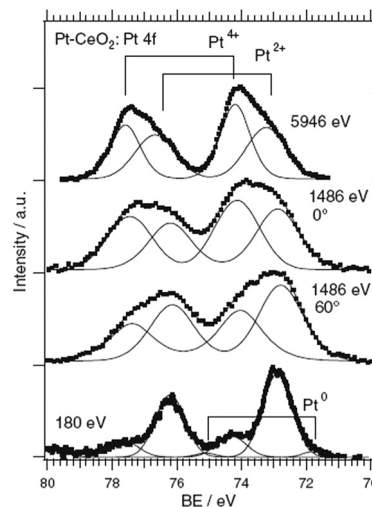


Fig. 2 Pt 4f spectra obtained from the Pt–CeO₂ film by means of SRPES ($h\nu = 180$ eV, $\alpha = 0^\circ$), AR XPS ($h\nu = 1486$ eV, $\alpha = 60^\circ$ and 0°), and HAXPS ($h\nu = 5946$ eV, $\alpha = 0^\circ$). Reproduced with permission from ref. 39, Copyright 2009, The Electrochemical Society.

metals.^{40,49,50} The ratio between the Pt²⁺, Pt⁴⁺, and Pt⁰ contributions in the Pt 4f is a function of the sampling depth. Accordingly, Pt⁴⁺ species are located mostly in the bulk while Pt²⁺ and Pt⁰ reside at the surface.

The relationship between the performance of the Pt–CeO₂ catalyst supported on n-GDL and the abundance of Pt²⁺, Pt⁴⁺, and Pt⁰ species can be derived from the comparison of the SPs (Table 1) and the shape of the corresponding Pt 4f spectra (Fig. 3a) as a function of Pt loading.

Clearly, the SP is directly related to the Pt²⁺/Pt⁴⁺ ratio in the films. The highest SP was obtained at the highest Pt²⁺/Pt⁴⁺ ratio corresponding to the lowest Pt loading ($0.6 \mu\text{g cm}^{-2}$). However, the PD shows a rather non-linear dependence on the Pt²⁺/Pt⁴⁺ ratio. The maximum PD was achieved at moderate Pt loading ($2 \mu\text{g cm}^{-2}$) and at a lower Pt²⁺/Pt⁴⁺ ratio. Note that a further increase of the Pt loading yielded a lower Pt²⁺/Pt⁴⁺ ratio and resulted in a substantial decrease of the PD (Table 1). The analysis of the Pt–CeO₂ catalyst after running several FC cycles revealed conversion of the Pt⁴⁺ species to Pt²⁺ and Pt⁰ (Fig. 3b). As a result, both Pt–CeO₂ films with Pt loadings 0.6 and $2 \mu\text{g cm}^{-2}$ contain Pt²⁺ species, exclusively. Despite the slightly lower SP, the Pt–CeO₂ film with

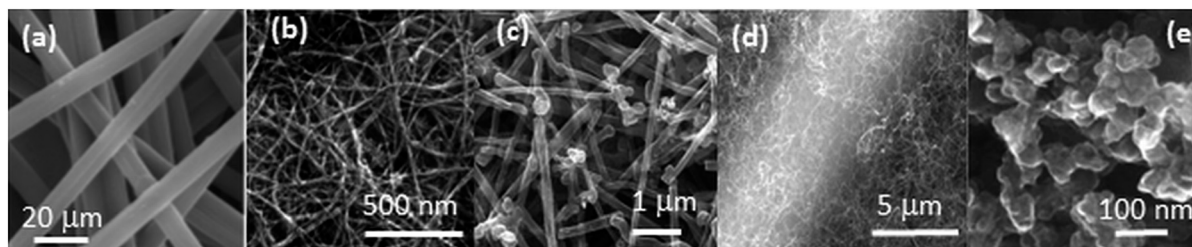


Fig. 1 SEM images of bare (a) and DWCNT (b), MWCNT (c), CVD-CNT (d) coated GDL, and n-GDL (e). (a) Reprinted from ref. 45 with permission from John Wiley & Sons, Inc. Copyright 2016 by John Wiley & Sons, Inc. (b) Adapted from ref. 40, Copyright 2009, The Electrochemical Society. (c) Reprinted from ref. 46 with permission from John Wiley & Sons, Inc. Copyright 2010 by John Wiley & Sons, Inc. (d) Reprinted from ref. 47, Copyright 2015, with permission from Elsevier.



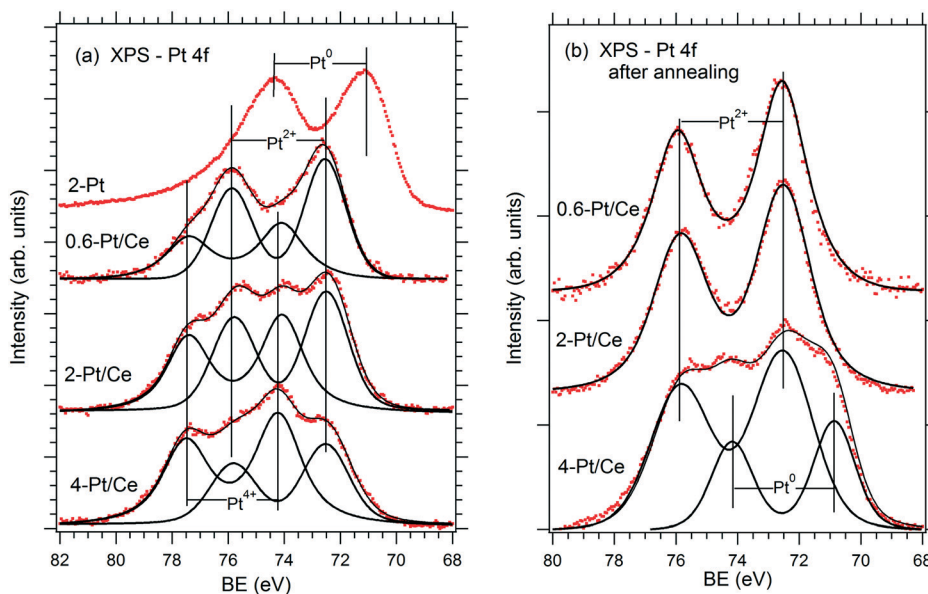


Fig. 3 Pt 4f spectra obtained from the reference Pt film ($2 \mu\text{g Pt cm}^{-2}$, 2-Pt) and Pt-CeO₂ films supported on n-GDL with Pt loadings of 0.6 (0.6-Pt/Ce), 2 (2-Pt/Ce) and 4 (4-Pt/Ce) $\mu\text{g Pt cm}^{-2}$ before (a) and after annealing under PEMFC operating conditions (b). Reprinted from ref. 41, Copyright 2016, with permission from Elsevier.

a Pt loading of $2 \mu\text{g cm}^{-2}$ shows a significantly higher PD with respect to the film with a Pt loading of $0.6 \mu\text{g cm}^{-2}$. Under these conditions, the PD increases with the density of Pt²⁺ species. From the perspective of volumetric power densities, the Pt-CeO₂ catalyst with a Pt loading of $2 \mu\text{g cm}^{-2}$ is therefore the most suitable for a compact design of the FC.

The presence of metallic Pt⁰ deteriorated the catalyst performance (see Fig. 3 and Table 1 for a Pt loading of $4 \mu\text{g cm}^{-2}$). It follows that the excellent performance of the Pt-CeO₂ catalyst is associated with the high density and the enhanced stability of Pt²⁺ species.

2.3. Parameters controlling proton exchange membrane fuel cell performance

The main parameter that controls the abundance of Pt²⁺ species on the Pt-CeO₂ catalyst is the morphology of the support. The SEM images obtained from Pt-CeO₂ deposited on CNT-coated GDL and n-GDL are shown in Fig. 4. The mor-

phology of the Pt-CeO₂ anode catalyst corresponds to a porous columnar structure⁴² that varies as a function of the support with respect to the width and height of the crystallites. The degree of nanostructuring increases for films deposited on CNT-coated GDLs (Fig. 4a-c) and n-GDLs (Fig. 4d) with respect to the bare GDL.

The parameters controlling the growth of Pt-CeO₂ films have been systematically investigated with respect to the microstructure of the carbon films, the pressure and the temperature. It was found that the major process giving rise to the porous structure of the Pt-CeO₂ films is associated with the etching of the carbon films by oxygen plasma.⁵²⁻⁵⁵ The corresponding mechanism involves the formation of nucleation centers, *e.g.* Pt-CeO_x or Ce-C particles which mask the carbon substrate partially and, thus, prevent etching.^{53,55} The density and the mobility of these nucleation centers determine the width of the columns. Additional parameters are the deposition rate⁵⁵ and thickness⁵⁶ of the Pt-CeO₂ films, temperature, and the composition of the reactive gas.^{54,55}

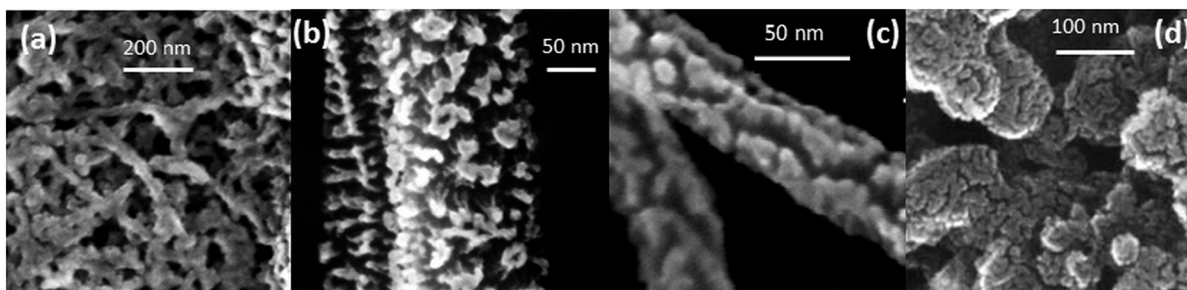


Fig. 4 SEM images of the Pt-CeO₂ anode catalyst deposited on DWCNT (a), MWCNT (b), and CVD-CNT (c) coated GDL, and n-GDL (d). (a) Adapted with permission from ref. 40, Copyright 2009, The Electrochemical Society. (b) Reprinted with permission from ref. 51, Copyright 2012, Inderscience Enterprises Ltd. (d) Reprinted from ref. 47, Copyright 2015, with permission from Elsevier.



2.4. Preparation of the nanostructured Pt–CeO₂ films

Besides radio frequency magnetron sputtering,^{21,39–44} the nanostructured Pt–CeO₂ films can be prepared by a variety of other techniques including pulsed laser deposition⁵⁷ and chemical vapor deposition techniques.^{58–61} The preparation of nanostructured Pt–CeO₂ films by means of physical vapor co-deposition of Pt and Ce in an oxygen atmosphere requires low deposition temperature.²¹ For some methods, *e.g.* pulsed laser deposition, the degree of the nanostructuring of the Pt–CeO₂ films depends on the morphology, *i.e.* roughness of the support.⁵⁷

2.5. Identification of the surface sites on the nanostructured CeO₂ support

The degree of the nanostructuring of the Pt–CeO₂ films determines the density of the Pt²⁺ species and is closely related to the number of Ce³⁺ cations.^{42,52} For instance, the Pt²⁺/Pt⁴⁺ ratio is much higher in the Pt–CeO₂ films deposited under glancing angle (GLAD) conditions.⁴⁶ In contrast, normal deposition (ND) yielded mainly Pt⁴⁺ species accompanied by a low number of Ce³⁺ cations.⁶² It was found that Ce³⁺ sites are located predominantly at the surface steps and edges of the nanostructured Pt–CeO₂ films.⁴²

The surface of the Pt–CeO₂ films was investigated by means of high-resolution transmission electron microscopy (HRTEM). There, the Pt–CeO₂ nanoparticles terminated by the {111} and {100} facets were identified in the Pt–CeO₂ films supported on CNT-coated GDL,^{42,43} n-GDL,^{41,44} and silicon wafers.^{42,63} Typical HRTEM images obtained from the Pt–CeO₂ films supported on n-GDL are shown in Fig. 5 for two different Pt loadings.

The structure of the Pt–CeO₂ films is identical to the structure of bare CeO₂ films and does not depend on the choice of the preparation technique. For instance, nanoparticles of CeO₂ and Pt–CeO₂ terminated predominantly by the {111} and {100} facets have been identified in films prepared by magnetron sputtering,^{56,64} chemical vapor deposition,⁶⁰ and pulsed laser deposition.⁵⁷ It follows that with respect to the nature of Pt²⁺ surface species, either the {111} or {100} facets or edge and step sites connecting

these facets provide the adsorption sites that anchor atomically dispersed Pt atoms.

3. Stability of single atom catalysts

3.1. Anchoring noble metal atoms at surface sites of nanostructured ceria

The capacity of the nanostructured ceria to anchor atomically dispersed noble metals at the {111} and {100} facets was analyzed by means of density functional calculations.²¹ A Ce₄₀O₈₀ nanoparticle model^{65,66} featuring a truncated octahedral shape with O-terminated {111} and very small {100} facets was identified as a representative model for nanostructured ceria.⁶⁷ This particle structure, displayed in Fig. 6, emerged from an exhaustive global optimization search.^{65,66,68} The polar (100) surface corresponding to the {100} nanofacets of the ceria particles is known to be less stable than the (111) surface of ceria.⁶⁹ Each of these {100} nanofacets is terminated by four surface O²⁻ ions in a square planar arrangement with O–O distances of 315–320 pm forming a so-called O₄ pocket.

The binding of Pt atoms adsorbed on the regular CeO₂(111) surface in the form of Pt⁰ or Pt⁺ is associated with a low diffusion barrier along the surface.⁷¹ The formation of Pt₂ dimers is therefore kinetically facile and strongly exothermic (by 369 kJ mol⁻¹)⁷² giving rise to rapid nucleation of Pt particles.⁷³ In turn, the oxidation state and adsorption energy of Pt atoms on the surface of ceria nanoparticles drastically depend on the local structure of the adsorption sites. A moderately strong binding of atomic Pt on edge sites between {111} facets of the Ce₄₀O₈₀ nanoparticle yields Pt⁰ and Pt⁺ species with the adsorption energies of –273 and –303 kJ mol⁻¹, respectively.²¹ Similar adsorption energies were calculated for Pt⁰ (–256 kJ mol⁻¹) and Pt⁺ species (–226 kJ mol⁻¹) on the regular CeO₂(111) surface.⁷¹ This adsorption strength of atomic Pt is typical also for regular surfaces of more inert non-reducible metal oxides, such as MgO(001).⁷⁴ In contrast, the adsorption of Pt atoms on the {100} nanofacet yields Pt²⁺ species with an extraordinarily large binding energy of –678 kJ mol⁻¹,²¹ which is ~100 kJ mol⁻¹ more in magnitude than that for Pt atoms anchored to neutral oxygen vacancies of the MgO(001) surface.⁷⁵ In these calculations the formal oxidation state of the adsorbed Pt atom was determined by the number of Ce³⁺ centers formed *via* electron transfer into the 4f orbitals of Ce⁴⁺. Consequently, the formation of Pt⁺ or Pt²⁺ species is accompanied by the reduction of one or two Ce⁴⁺ cations to Ce³⁺, respectively (Fig. 6).

Strong interactions of metal atoms with oxide surfaces are indicative of surface coordination compounds, in which the support acts as a polydentate ligand.⁷⁶ Indeed, the O₄ site provides an ideal coordination environment to host the Pt²⁺ (d⁸) species yielding a square planar PtO₄ moiety.⁷⁷ Remarkably, the adsorption energy of the anchored Pt²⁺ ion exceeds in magnitude the cohesive energy of bulk Pt (–564 kJ mol⁻¹).⁷⁸ As a consequence, Pt²⁺ species are thermodynamically stable against sintering. The structure of the resulting PtO₄ moiety is consistent with the interatomic distances and

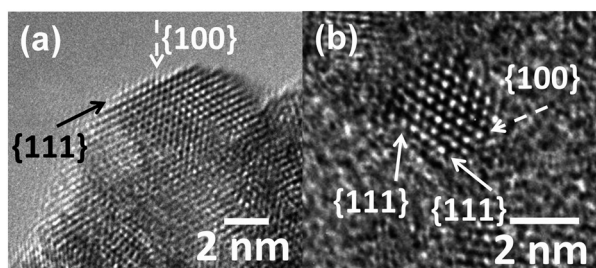


Fig. 5 HRTEM images of the Pt–CeO₂ films supported on n-GDL with Pt loadings of 2 (a) and 4 (b) μg cm⁻². The arrows show the location of the {111} and {100} facets. Reprinted from ref. 41, Copyright 2016, with permission from Elsevier.



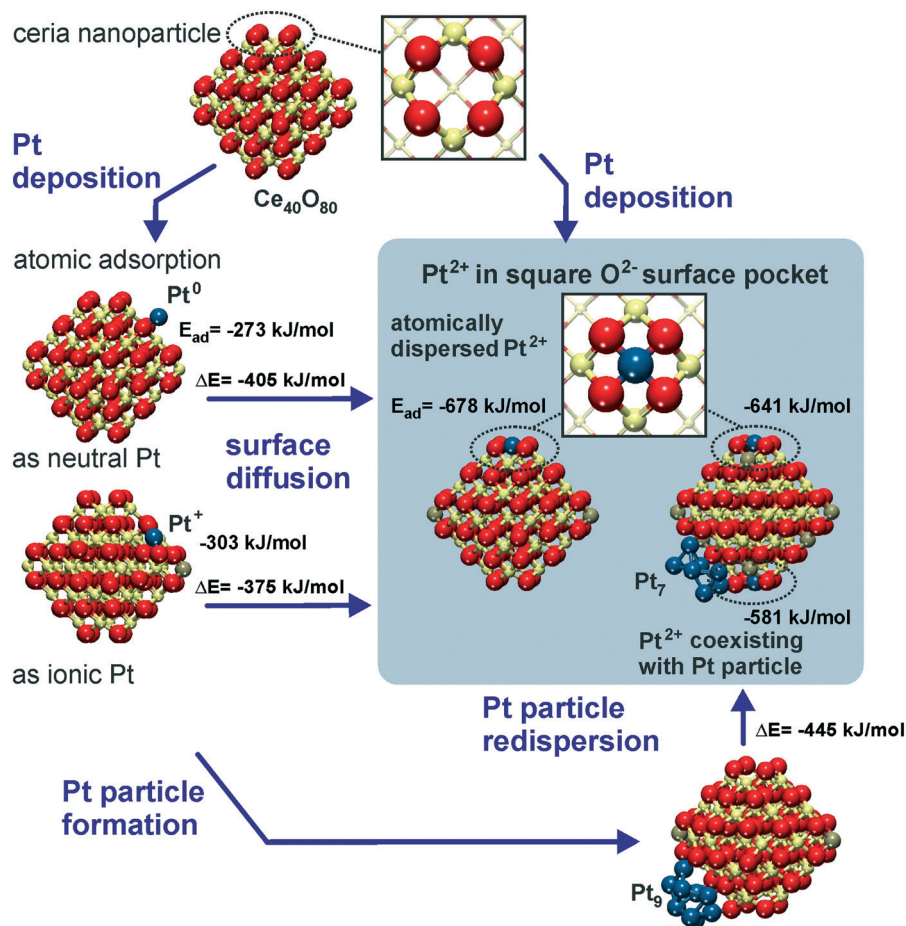


Fig. 6 Structure and energetics of various anchored Pt species on ceria nanoparticles obtained from density functional calculations. Pt^{2+} is strongly bound to square O_4 pockets at the $\{100\}$ nanofacets of the particles. Color coding of atoms: red O, beige Ce^{4+} , brown Ce^{3+} , and blue Pt. Reproduced with permission from ref. 70, Copyright 2016, Springer International Publishing AG.

coordination numbers of Pt atoms determined by means of extended X-ray absorption fine structure (EXAFS) studies on real Pt/ceria catalysts.⁷⁹

With respect to sintering and re-dispersion processes, the DFT calculations suggest that the O_4 sites are even able to abstract Pt atoms from supported Pt particles²¹ (Fig. 6). This pathway was calculated to be exothermic for Pt_9 and Pt_8 particles on $\text{Ce}_{40}\text{O}_{80}$, where the migration of Pt atoms from the metal particle to the $\{100\}$ nanofacets of ceria leads to their transformation into strongly bound Pt^{2+} species.

Furthermore, the anchored Pt^{2+} species do not serve as a stable nucleation site for a second Pt atom. Accordingly, no local minimum corresponding to a $\text{Pt}-\text{Pt}^{2+}/\text{O}_4$ moiety was found and during geometry optimization a supported Pt_2 dimer dissociated into a $\text{Pt}^{2+}/\text{O}_4$ complex and a neutral Pt atom adsorbed nearby. This finding implies that the anchored Pt^{2+} species can actually coexist with Pt particles without being buried by excess Pt.

A very similar local PtO_4 structure emerges upon Pt adsorption at the steps of the $\text{CeO}_2(111)$ surface.²² The thermodynamics of segregation and the corresponding atomic and electronic structures of Pt on stepped $\text{CeO}_2(111)$ were investi-

gated by density functional calculations.²² Different adsorption sites were considered (Fig. 7) including oxygen vacancies (a), regular sites (b) and Pt clusters (c) on the $\text{CeO}_2(111)$ terrace, and the two low-energy one-monolayer-high steps labeled as step I (d, f) and step II (e, g) on stoichiometric (denoted S) and non-stoichiometric (with excess of oxygen, denoted O) surfaces. Structures I and II represent the preferred types of steps at the $\text{CeO}_2(111)$ surfaces at temperatures below 1000 K.^{80,81}

Pt adsorption at step I-S yields Pt^{2+} species coordinated by four lattice O atoms in a characteristic PtO_4 planar unit (Fig. 7d) similar to that identified on the $\{100\}$ facets of the $\text{Ce}_{40}\text{O}_{80}$ nanoparticle.²¹ Consequently, two Ce^{3+} centers are formed in the proximity of the Pt^{2+} (Fig. 7d). The resulting adsorption energy of Pt^{2+} is -5.0 eV . In contrast, the different atomic structure of the step II-S edge prevents the formation of the PtO_4 moiety and yields a weakly oxidized Pt^{4+} species accompanied by the emergence of one Ce^{3+} center (Fig. 7e).

The presence of excess oxygen at the steps triggers restructuring and yields Pt^{2+} in the characteristic square-planar PtO_4 surface arrangement. At step I, the excess oxygen leads to the oxidation of Ce^{3+} centers to Ce^{4+} while Pt^{2+}



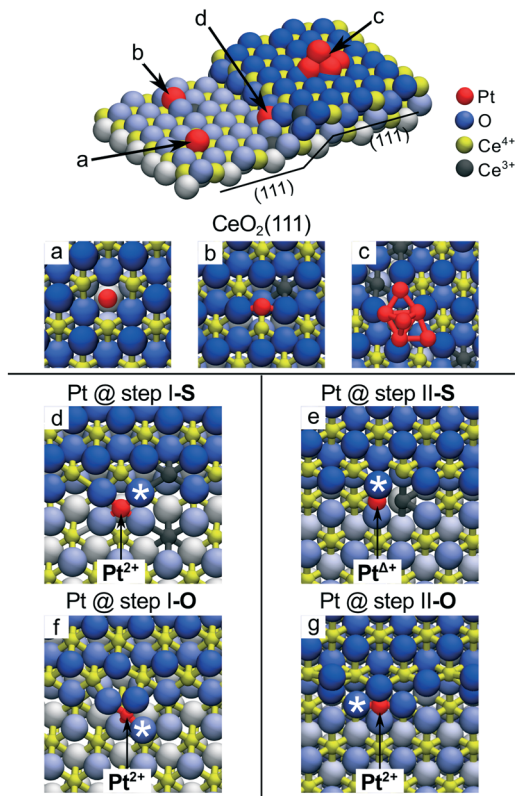


Fig. 7 Pt adsorption sites on the stepped CeO₂(111) surface. (a) Pt adatom in the surface O vacancy, (b) on the stoichiometric CeO₂(111) terrace and (c) supported Pt₆ cluster. (d) Pt adatom at the stoichiometric step I (step I-S) and (e) at the stoichiometric step II (step II-S). (f) Pt adatom at step I with excess O (step I-O) and (g) at step II with excess O (step II-O). Reproduced with permission from ref. 22, Copyright 2016, Nature Publishing Group.

retains its oxidation state. At step II, the oxidation of the Pt⁴⁺ species to Pt²⁺ is accompanied by the oxidation of one Ce³⁺ to Ce⁴⁺. Remarkably, the resulting binding energies of Pt²⁺ species in excess of oxygen are ~ 1.6 eV higher (*i.e.* more strongly bound) with respect to that calculated for Pt adsorption on stoichiometric step sites. These results demonstrate that stable PtO₄ moieties with Pt²⁺ species can be formed at step sites of ceria upon oxidation of Pt⁰ to Pt²⁺ and that the formation of such Pt²⁺ species does not necessarily involve the formation of Ce³⁺ ions.²²

The calculations predict the preferential adsorption of Pt atoms at steps I and II regardless of their local geometry and stoichiometry. This prediction is consistent with the higher density of Pt²⁺ species on ceria surfaces featuring higher step density. Furthermore, it implies that the same square-planar arrangement of PtO₄ moieties is formed on different nanostructured ceria supports including stepped surfaces and nanoparticles.

The capacity of the {100} sites on nanostructured ceria to anchor atomically dispersed Pt suggests that other transition metals could interact with the O₄ site of Ce₄₀O₈₀ in a similar way. In this respect, the adsorption of metal atoms (M) including the 3d (Fe, Co, Ni, Cu), 4d (Ru, Rh, Pd, Ag) and 5d

(Os, Ir, Pt, Au) metals of groups 8–10 was investigated.⁸² The calculated adsorption energies are shown in Fig. 8 in comparison with the calculated adsorption energy of an edge metal atom of a M₇₉ nanoparticle of the corresponding metal.

The adsorption of the metal atom M at the {100} site results in the oxidation of the adsorbed atom, accompanied by the reduction of Ce⁴⁺ cations to Ce³⁺. In all cases, the adsorption of the metal atom on the O₄ site of the ceria nanoparticle is stronger than the binding of edge metal atoms in the M₇₉ models. Interestingly, all calculated oxidation states of metal atoms of groups 8–11 are also observed experimentally in metal–ceria systems.^{21,26,65,79,83–97} However, only the metals of group 10 all yield cationic species of the same oxidation state, *i.e.* Pt²⁺, Pd²⁺, and Ni²⁺. This similarity is associated with the d⁸ electronic configuration of the M²⁺ cations in the group 10 metals adsorbed in the square planar arrangement. The metals of group 11 adsorbed at the {100} site are found in multiple oxidation states. For instance, Cu²⁺ in the square planar coordination is just slightly more stable than the Cu⁺ in a distorted linear arrangement with two short and two long Cu–O distances. Similarly, atomic Ag was found as Ag⁺ or Ag³⁺ species. Atomic Au can also be adsorbed as Au⁺ and Au³⁺, with the latter being more stable.

Clear trends emerge along the rows and groups of the periodic table (Fig. 8). Binding energies of M both on the ceria nanoparticle and in the M₇₉ nanoparticles generally decrease

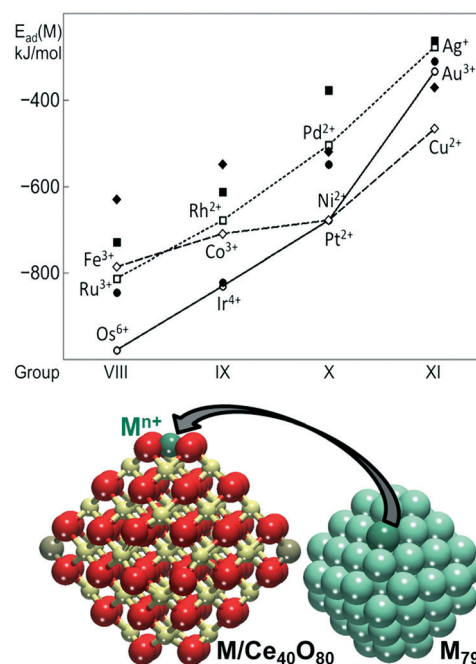


Fig. 8 Adsorption energies of atomically dispersed 3d (diamonds), 4d (squares), and 5d (circles) metals (M) on the {100} facet of the Ce₄₀O₈₀ nanoparticle (empty symbols, connected for guiding the eye) and binding energies of these atoms in edge positions of the M₇₉ NP (filled symbols). The corresponding M–Ce₄₀O₈₀ and M₇₉ structures are also shown. Yellow, brown and red spheres represent Ce⁴⁺, Ce³⁺ and O²⁻ ions, respectively. Adapted from ref. 82 with permission from the Royal Society of Chemistry.



in magnitude, when moving from the left to the right of the period. This indicates that metals with less occupied d bands form stronger metal–metal bonds and also bind more strongly to the oxide support. With the exception of Au, 5d metals form the strongest bonds with the O₄ site of ceria, whereas 4d metals form the weakest bonds, with the exception of Ru. Regardless of the overall strength of the interaction of the different metal atoms with the ceria nanoparticle, the resulting adsorption energies indicate that the {100} sites are able to stabilize atomically dispersed species of a wide range of metals.

3.2. Stability as a function of noble metal loading

The stability of the ceria supported single atom catalysts is a function of numerous parameters including the metal loading,⁹⁸ the structural and chemical environment of the adsorption site,²² the oxidation state of the substrate, *i.e.* the density of Ce³⁺ centers²¹ and oxygen vacancies,^{89,99} and the ambient atmosphere.⁸⁹

For instance, the adsorption energy of the Pt²⁺ species at {100} sites decreases from 7.02 eV to 6.59 eV when the Pt coverage increases from one to four Pt²⁺ species per Ce₄₀O₈₀ NP due to the increase of Ce³⁺ density.⁹⁹ With respect to the limited capacity of the Ce₄₀O₈₀ NP (only 4 Pt atoms can be adsorbed in the form of Pt²⁺ since there are only four {100} nanofacets each exposing one O₄ site), the stepped CeO₂(111) surface allows a higher density of Pt²⁺ species to be obtained by assembling PtO₄ units along the steps in close proximity to each other (see Fig. 9).

The maximum coverage of the stable Pt²⁺ species at step I-S is 2/3 (Fig. 9a). A further increase of the Pt coverage to 3/3 (1 Pt atom per 1 Ce step edge atom) triggers the nucleation of metallic Pt clusters due to the large strain built up in the long row of interconnected PtO₄ units at the I-S step (Fig. 9b). On step II-S, reduction of Pt²⁺ species gives rise to Pt dimers already at coverages exceeding 1/3 (Fig. 9c). In summary, on samples with stoichiometric steps I and II, the calculations predict low Pt²⁺ coverage at the steps (<33% of the step-edge sites). In contrast, the calculated maximum coverage of Pt²⁺ at the O-rich steps I-O and II-O is 3/3 (Fig. 9e and g), as interconnected assemblies of PtO₄ units fit to the periodicities of both steps I and II (Fig. 7f, g and 9d–g).

The stability of PtO₄ moieties was investigated more quantitatively by evaluating the Gibbs free energy of Pt²⁺ adsorbed at the {100} site of a Ce₂₁O₄₂ particle as a function of ambient temperature and partial pressure of oxygen.⁸⁹ The relative stability of oxygen vacancies or excess oxygen species evolves progressively with changing pressure and temperature (Fig. 10). Under oxidizing conditions, the adsorption of one oxygen atom in the proximity of the Pt²⁺ species is favored and leads to the re-oxidation of two Ce³⁺ centers to Ce⁴⁺. Under reducing conditions, the adsorption energy of the Pt²⁺ species at the {100} site decreases with increasing number of oxygen vacancies. At higher temperature, the onset of vacancy formation shifts to higher oxygen pressure and approaches the onset of oxygen adsorption, thus narrowing the region of the Pt²⁺ stability.

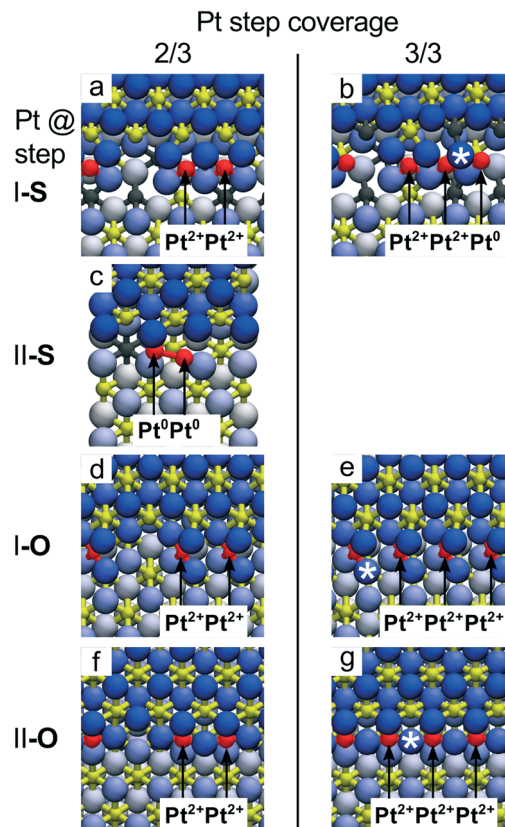


Fig. 9 Capacity of the CeO₂(111) step edges to accommodate Pt²⁺ ions obtained from DFT calculations. Calculated top views of Pt atoms bound to steps I-S (a and b), II-S (c), I-O (d and e) and II-O (f and g) for Pt step coverage 2/3 (a, c, d and f) and 1 (b, e and g). At step I-S, the limiting coverage of Pt²⁺ is 2/3 (a), additional Pt attaches to the step edge as Pt⁰ (b). At step II-S, the Pt²⁺ coverage is 0. Pt atoms attach as weakly ionized Pt¹⁺ and readily form metallic dimers (c) and clusters. On both steps I-O and II-O, excess oxygen can stabilize ionic Pt²⁺ at step edges as single ions appearing isolated or in groups up to 100% step coverage (d–g). Reproduced with permission from ref. 22, Copyright 2016, Nature Publishing Group.

The theoretically predicted behavior of atomically dispersed Pt was examined using appropriate model systems based on well-defined surfaces prepared under UHV conditions.^{21,22,98} In particular, Pt²⁺ species were prepared either on the surface of nanostructured ceria films²¹ or on stepped CeO₂(111) surfaces.²² The first approach employed co-deposition of Pt and Ce metal in an oxygen atmosphere onto a well-ordered CeO₂(111) buffer layer at low temperature. The use of a 1.5 nm thick buffer layer was necessary to minimize the influence of the Cu(111) substrate on the structure of the Pt–CeO₂ film.¹⁰⁰ The procedure yielded three-dimensional Pt–CeO₂ particles with an average diameter of approximately 3 nm and a typical height of around 0.4 nm (see Fig. 11a and b).²¹

After annealing to 700 K, some aggregates reveal a faceted shape, suggesting an epitaxial relationship between the supported Pt–CeO₂ nanoparticles and the CeO₂(111) buffer layer. The chemical state of Pt in the Pt–CeO₂ film was investigated by means of SRPES with high surface sensitivity. Two



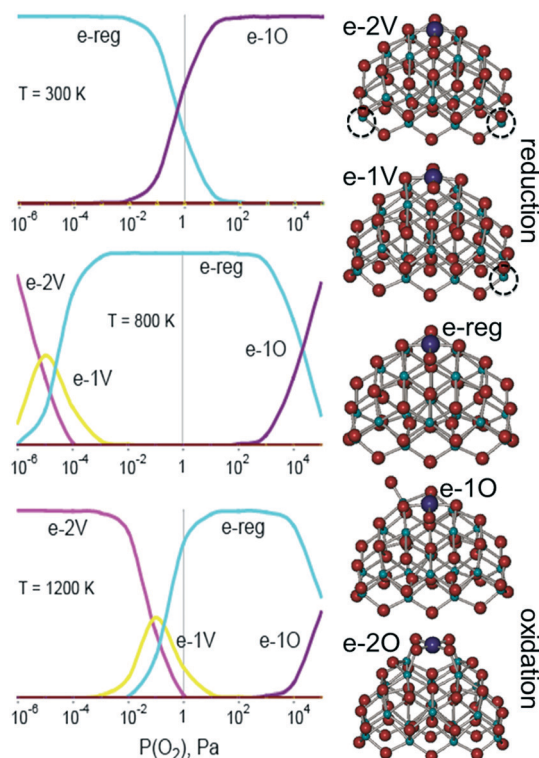


Fig. 10 Left: results from a general thermodynamic model including a series consisting of the most stable structures of reduced (e-2v and e-1v), stoichiometric (e-reg), and oxidized (e-1O and e-2O) $\text{PtCe}_{21}\text{O}_{42}$ at three different temperatures. The vertical axis corresponds to the relative concentration of the species ranging from 0 to 100%. Right: structures for the different states considered (dashed circles correspond to the positions of oxygen vacancies). Adapted from ref. 89 with permission from the PCCP Owner Societies.

types of species associated with Pt^{2+} and Pt^{4+} ions were identified in the Pt 4f spectra obtained from the Pt– CeO_2 films as deposited (Fig. 11c and d). In the limit of low Pt loading, the Pt^{2+} species in the Pt– CeO_2 films demonstrate exceptional thermal stability upon annealing up to 750 K (see Fig. 11c). At the same time, less stable Pt^{4+} species were readily converted to Pt^{2+} . The critical role of providing the appropriate number of surface sites to anchor Pt becomes evident when the Pt– CeO_2 films with higher Pt loading were annealed (Fig. 11d). Consequently, the Pt atoms anchored at less favorable sites are bound too weakly to resist sintering to metallic particles.

In line with the density functional calculations⁸² described in section 3.1, Pd^{2+} and Ni^{2+} species were also anchored on the surface of nanostructured ceria following a similar experimental approach. These films were characterized under similar conditions to the Pt– CeO_2 films by means of SRPES (see Fig. 12 and 13, respectively).

The formation of both Pd^{2+} and Ni^{2+} species was accompanied by additional oxide phases, namely PdO and NiO, that resulted from the lack of a sufficient number of {100} sites on the nanostructured ceria films (see Fig. 12 and 13). Note that the formation and decomposition of the NiO phase can be identified by the presence of two satellite features, *i.e.* I and II in the spectra of Ni 2p in Fig. 13. The abundance of the addi-

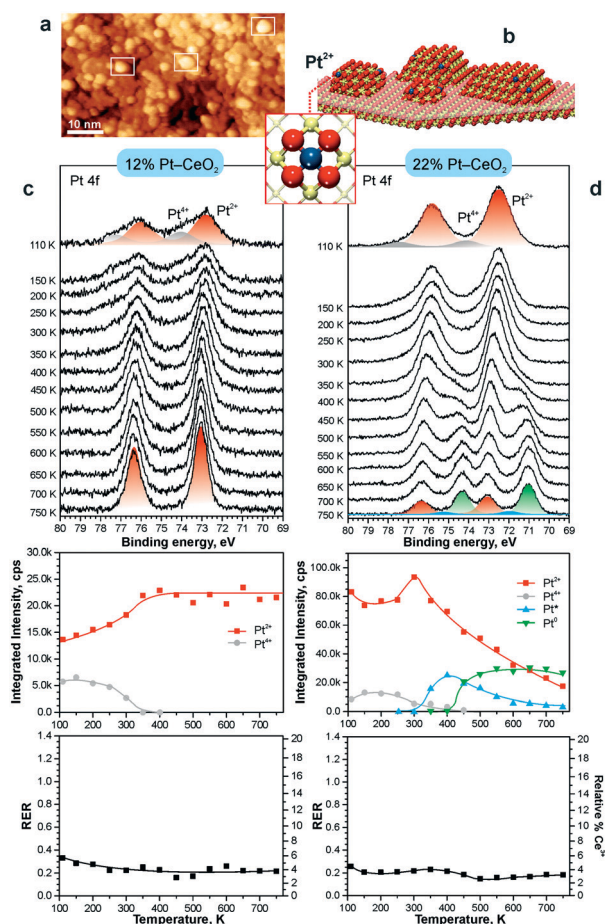


Fig. 11 a) STM image obtained from Pt– CeO_2 nanoparticles supported on the $\text{CeO}_2(111)$ buffer layer after annealing to 700 K. White rectangles outline faceted Pt– CeO_2 particles. b) Schematic structure of the model catalyst. The height of Pt– CeO_2 particles is amplified for illustrative purposes. c and d) Pt 4f spectra, the integrated intensities of the surface species, and RER as a function of annealing temperature at a Pt loading of 12% (c) and 22% (d) in the volume of Pt– CeO_2 films. The nominal thickness of the Pt– CeO_2 films is 0.3 nm. Pt 4f spectra were obtained with a photon energy of 180 eV. (a and b) reproduced from ref. 21 with permission from John Wiley & Sons, Inc. Copyright © 2016 by John Wiley & Sons, Inc. (c and d) Adapted with permission from ref. 98, Copyright 2016, American Chemical Society.

tional phases, *e.g.* oxide and metal particles is a function of metal loading. Thus, in the limit of the low metal loading, the most stable Pd^{2+} and Ni^{2+} species can be isolated from the less stable oxide phases by annealing in UHV (see Fig. 12a and b and 13a and b). Both PdO and NiO phases decompose *via* dissolution in the $\text{CeO}_2(111)$ buffer layer upon annealing.

Higher Pd loading in the films results in a low stability of atomically dispersed Pd^{2+} species (see Fig. 12d and e). Similar to the Pt^{2+} , atomically dispersed Pd^{2+} species are converted into metallic particles upon annealing in UHV.

However, the characteristic difference between the Pt^{2+} and Pd^{2+} species is that Pt^{2+} is preferentially stabilized at the oxygen pockets at the surface, whereas Pd can be stabilized both at the surface and in the bulk in a similar square planar arrangement.⁹⁴ This property facilitates the diffusion of Pd^{2+} species into the bulk upon annealing. In sharp contrast to



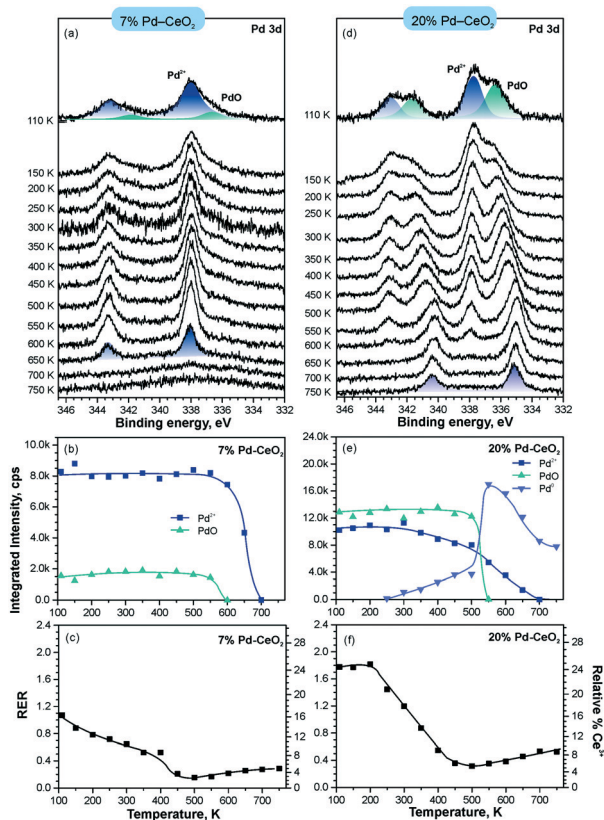


Fig. 12 Stability of 7% Pd-CeO₂ (a–c) and 20% Pd-CeO₂ (d–f) upon annealing in UHV. Pd 3d spectra (a and d) obtained with $h\nu = 410$ eV. The integrated intensities of the surface species (b and e) and RER (c and f) on 7% Pd-CeO₂ (a–c) and 20% Pd-CeO₂ (d–f) as a function of temperature. Adapted with permission from ref. 98, Copyright 2016, American Chemical Society.

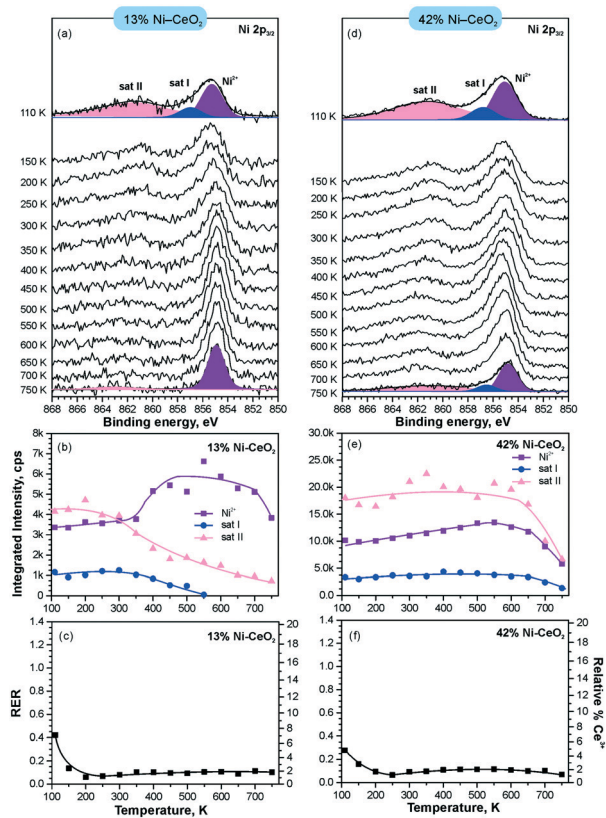


Fig. 13 Stability of 13% Ni-CeO₂ (a–c) and 42% Ni-CeO₂ (d–f) upon annealing in UHV. Ni 2p_{3/2} spectra (a and d) obtained with $h\nu = 1000$ eV. The integrated intensities of the surface species (b and e) and RER (c and f) on 13% Ni-CeO₂ (a–c) and 42% Ni-CeO₂ (d–f) as a function of temperature. Adapted with permission from ref. 98, Copyright 2016, American Chemical Society.

both Pt-CeO₂ and Pd-CeO₂, formation of metallic Ni particles was not observed in Ni-CeO₂, even at high metal loading. This observation is in agreement with the calculated data suggesting that the formation of metallic Ni from ionic Ni²⁺ is strongly disfavored energetically.

Interestingly, the atomically dispersed metals of group 10 yielded different concentrations of Ce³⁺ centers. This again demonstrates the diversity of mechanisms involved in the anchoring. In contrast to the Pt-CeO₂ films, the Pd-CeO₂ films maintain a considerably higher degree of reduction even at low dopant concentrations.⁸ The decomposition of PdO and NiO phases during annealing results in re-oxidation of Ce³⁺.

An alternative approach to form Pt²⁺ species involves the deposition of Pt onto stepped CeO₂(111) surfaces in UHV. Under these conditions, the availability of the adsorption sites capable of anchoring Pt²⁺ species is the most critical parameter. The Pt deposition on CeO₂(111) surfaces with a low density of steps and surface oxygen vacancies (Fig. 14a) yields metallic Pt⁰ clusters in combination with ionic Pt²⁺ species (Fig. 14b and c).

Subsequent annealing of the films to 700 K increases the amount of Pt²⁺ species at the expense of Pt⁰. Interestingly, this process was not accompanied by the reduction of the CeO₂(111) surface. This implies the involvement of another oxidizing agent such as excess oxygen atoms. In the UHV environment,

one possible source of excess oxygen is water adsorbing in sub-monolayer amounts from the background and undergoing dissociation on reduced ceria and Pt/ceria substrates.^{101,102} In the real Pt-CeO₂ catalysts, excess O atoms may also be incorporated during the synthesis that proceeds in air.^{103,104}

Most importantly, it was found that the relative abundance of the Pt²⁺ and Pt⁰ species depends on the density of steps and the number of oxygen vacancies in the CeO₂(111) film. In particular, the amount of Pt²⁺ species scales linearly with the step density.

The presence of oxygen vacancies does not promote the dispersion of Pt²⁺ species but, instead, leads to the formation of small metallic particles (Fig. 14e and f). According to density functional calculations, oxygen vacancies are not the favorable sites for Pt²⁺ formation.^{22,89} Based on the analysis of the adsorption energies of Pt, formation of metallic clusters is preferred on reduced ceria nanoparticles.⁸⁹

3.3. Identification of the oxidation state of the supported catalyst by CO adsorption

The straightforward identification of Pt sites under reaction conditions is important to understand and tune the structure and stability of Pt-CeO₂ films. CO is commonly used as a



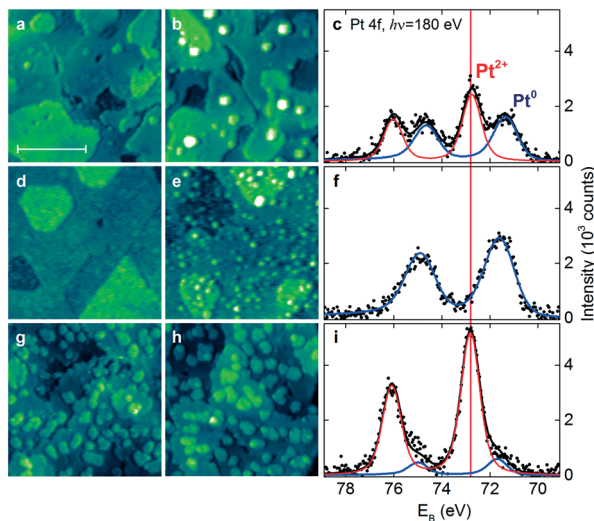


Fig. 14 STM images (a and b, d–e and g–h) and Pt 4f spectra (c, f and i) obtained from the stepped $\text{CeO}_2(111)$ surfaces with a low density of surface oxygen vacancies and steps (a–c), a high density of surface oxygen vacancies (d–f), and a high density of steps (g–i) before (a, d and g) and after (b, e and h) Pt deposition in UHV followed by annealing to 700 K. The size of the images is $45 \times 45 \text{ nm}^2$ and the scale bar is 20 nm. All STM images were obtained with a tunneling current of 25–75 pA and a sample bias voltage of 2.5–3.5 V. Pt 4f spectra were acquired with photon energy $h\nu = 180 \text{ eV}$. Reproduced with permission from ref. 22, Copyright 2016, Nature Publishing Group.

probe molecule in combination with infrared absorption spectroscopy (IRAS) for identification of adsorption sites in heterogeneous catalysts.^{105–108}

The factors that influence the vibrational frequency of CO adsorbed on Pt particles in Pt– CeO_2 systems include the local structure of the adsorption site (on-top, bridge, and hollow), the oxidation state and coordination number of the Pt adsorption center, the CO coverage, and the stoichiometry of the support.

For instance, the vibrational frequency of CO adsorbed in the on-top configuration on Pt surfaces increases as a function of the coordination number of the Pt atom. The effect was demonstrated for Pt(111) and unsupported Pt particles as well as for ceria-supported Pt particles of different sizes^{109,110} (Fig. 15).

As under-coordinated sites have characteristic IRAS fingerprints and the relative number of these sites depends on particle size, the trends derived from Fig. 15 allow estimating particle sizes on the basis of the measured CO vibrational frequencies (and *vice versa*).¹¹¹ For example, the size of Pt particles under electrochemical working conditions in Pt–ceria electrocatalysts was determined by comparing the experimental frequency shifts in IRAS with calculated shifts, both on free and ceria-supported Pt particles of varying size.¹⁰⁹

The calculated vibrational frequencies of CO adsorbed in the on-top configuration on Pt species with different oxidation states and coordination numbers of Pt atoms found in several structural Pt– CeO_2 models¹¹⁰ are summarized schematically in Fig. 16. A quick inspection of these data reveals a high degree of overlap between CO frequencies calculated

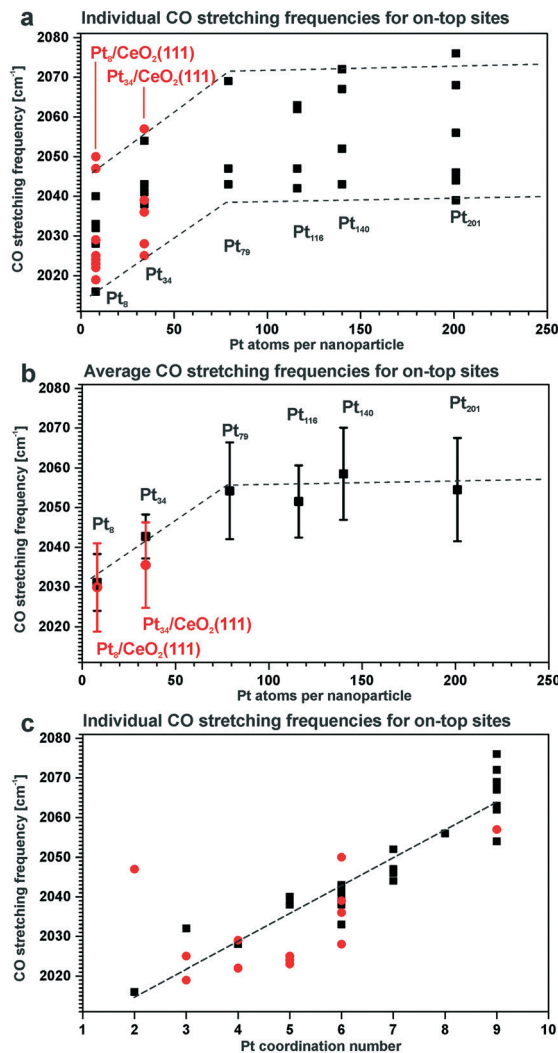


Fig. 15 Summary of the stretching frequencies calculated by DFT on different on-top sites of each Pt model particle (black, unsupported particles; red, supported particles): (a) CO stretching frequencies on all on-top sites of each particle and the corresponding tendency visualized by the dashed lines; (b) average on-top stretching frequencies and their standard deviation taking into account weights of all on-top sites in each particle; (c) correlation between the CO stretching frequency and the Pt coordination number with respect to neighboring Pt atoms. Reproduced with permission from ref. 109, Copyright 2016, American Chemical Society.

for the different structures. However, general trends indicate that the vibrational frequency of the CO molecule in the on-top configuration depends on the distance between Pt and the nearest oxygen atom of ceria at which the platinum atom is bound to, and that CO vibrational frequencies increase as a function of the oxidation state of atomically dispersed Pt on ceria.¹¹⁰ In addition, an increase in CO coverage leads to a widening of the CO frequency range for adsorption on supported Pt_8 clusters and to an increase of the frequency on Pt^{2+} species.

This scenario suggests that an assignment of platinum species based exclusively on the vibrational frequency might be misleading or inconclusive. When probing metal species



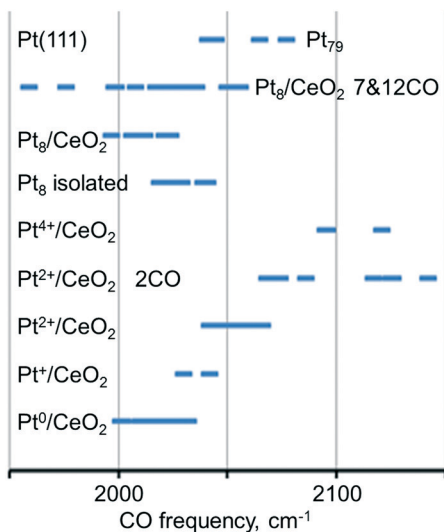


Fig. 16 Regions of the calculated C–O vibrational frequency for CO adsorbed on-top of neutral and cationic platinum species. Adapted from ref. 110 with permission from the PCCP Owner Societies.

with CO and IRAS, it is therefore necessary to monitor the coverage dependent changes in the spectra carefully. In addition, it appears that the adsorption energy of CO is a critical parameter. For example, the adsorption of CO on the Pt^{2+} cation of the PtO_4 moiety is rather weak (-0.61 eV) and also involves a significant reconstruction, suggesting that this is an activated and not very favorable process. Similarly, the vibrational frequency calculated for CO adsorption on Pt^{4+} comes from a rather unstable system, where excess O atoms have been added to obtain the Pt^{4+} state. This model is probably not very representative of sites one would find in experiments on Pt–ceria systems, but it is nevertheless useful as a tool to explore how the CO frequency changes with the oxidation state of Pt. CO does not adsorb under pertinent experimental conditions on Pt^{2+} and Pt^{4+} species based on the calculated stability of the resulting adsorption complexes.

As a complementary tool, SRPES allows direct determination of the oxidation state of Pt as well as the adsorption site of molecular surface species. The changes in the oxidation state of atomically dispersed Pt species caused by annealing in UHV were reliably identified from their Pt 4f spectra.¹¹² In particular, comparison of the intensities of the Pt 4f spectra before and after CO adsorption at 110 K indicated the presence of metallic Pt clusters which served as adsorption sites for CO.

The development of the Pt 4f spectra obtained from Pt– CeO_2 films with different Pt loadings are shown in Fig. 17. CO adsorption on the surfaces containing Pt^{2+} and Pt^{4+} exclusively does not attenuate the corresponding contributions (I–II) in the Pt 4f spectra. This suggests that neither Pt^{2+} nor Pt^{4+} serve as adsorption sites for CO in line with the calculated instability of CO on these sites. In contrast, small metallic Pt particles formed at high Pt loading (see Fig. 17c) readily adsorb CO. The emergence of metallic particles yielded distinct differences between Pt 4f spectra before (black) and after (green) CO adsorption (Fig. 17). In particu-

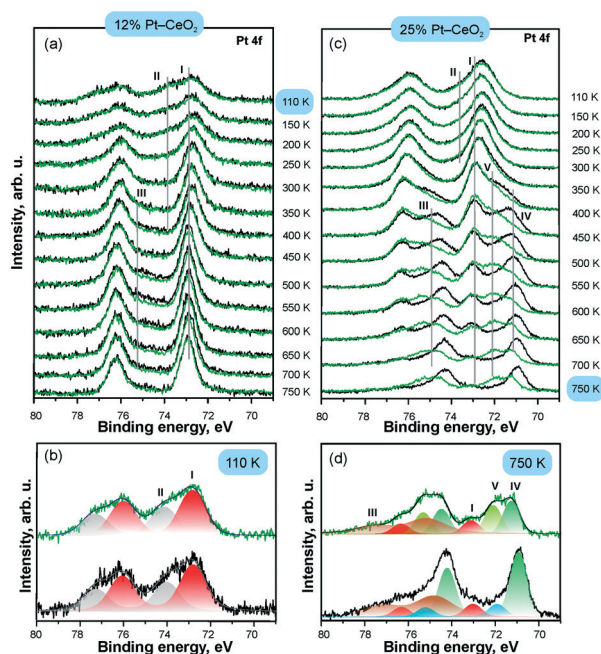


Fig. 17 Pt 4f spectra obtained from 12% Pt– CeO_2 (a and b) and 25% Pt– CeO_2 (c and d) films prepared at 110 K on $\text{CeO}_2(111)/\text{Cu}(111)$ annealed at different temperatures (black) and exposed to 50 L of CO at 110 K (green); the structure of the Pt 4f spectra on as-prepared 12% Pt– CeO_2 at 110 K (b) and 25% Pt– CeO_2 films annealed to 750 K (d); the components labelled I–V arise from Pt^{2+} , Pt^{4+} , Cu 3p, and metallic Pt particles (IV–V), respectively. Adapted from ref. 112 with permission from the PCCP Owner Societies.

lar, the contribution from metallic Pt (IV) in the Pt 4f spectra is attenuated and a new component (V) emerges in the Pt 4f spectra upon CO adsorption. Peak V is identified as a shifted component of peak IV, which is consistent with the core level shift expected upon CO adsorption on metallic Pt.

The development of the C 1s spectra upon CO adsorption on Pt– CeO_2 films with different Pt loadings is shown in Fig. 18 in comparison with the Pt-free nanostructured CeO_2 film. The C 1s spectra reveal the presence of CO-derived species on the three films. The dominant species formed on all surfaces are tridentate (A) and bidentate (B) carbonates. The emergence of the second tridentate carbonate (C) and carbonite (D) contributions indicates the presence of oxygen vacancies in the films. The formation of carbonite species is not associated with the formation of the tridentate carbonate species at oxygen vacancies (C) suggesting that the two species are formed at structurally different types of oxygen vacancies.

The emergence of a wide range of adsorbed molecular surface species including carbonates, carbonites, and formates, whose stability and spectroscopic signatures strongly depend on the structure and oxidation state of the adsorption site is in line with the results of DFT studies of CO adsorption on a $\text{Ce}_{21}\text{O}_{42}$ particle.¹¹³

The changes in the relative abundance of species A–D in Fig. 18 are consistent with restructuring of the films upon annealing. The morphological changes induced by the



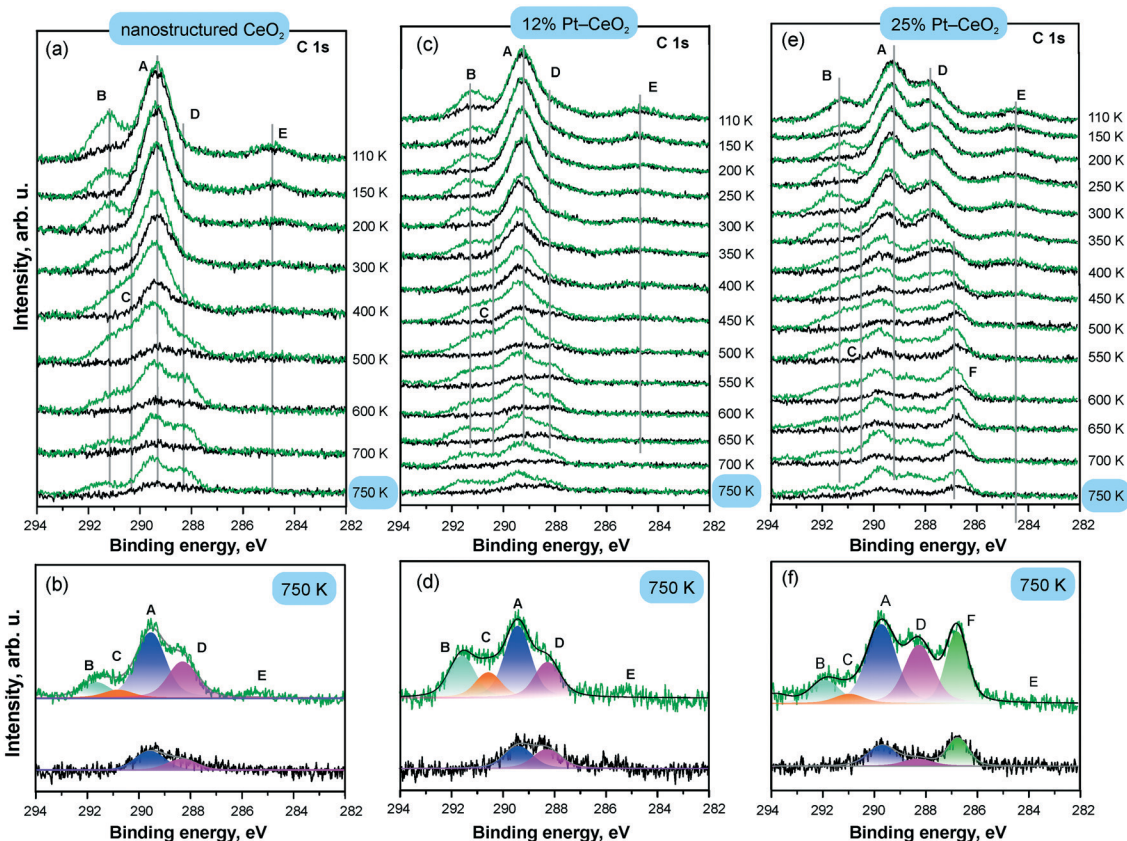


Fig. 18 C 1s spectra obtained from Pt-free nanostructured CeO₂ (a and b), 12% Pt–CeO₂ (c and d), and 25% Pt–CeO₂ (e and f) films prepared at 110 K on CeO₂(111)/Cu(111) annealed at different temperatures (black) and exposed to 50 L of CO at 110 K (green); the structure of the C 1s spectra obtained from Pt-free nanostructured CeO₂ (b), 12% Pt–CeO₂ (d), and 25% Pt–CeO₂ film (f) annealed to 750 K before (black) and after CO adsorption (green). The components labelled A–F arise from tridentate (A and C) and bidentate carbonates (B), carbonite (D), atomic carbon (E), and CO adsorption on metallic Pt particles (F). Adapted from ref. 112 with permission from the PCCP Owner Societies.

annealing of the 12% and 25% Pt–CeO₂ mixed oxides are similar to the changes of the Pt-free CeO₂ film and occur above 300 K. The observed reduction of Pt²⁺ to metallic Pt particles on the 25% Pt–CeO₂ film is also associated with restructuring causing a decrease in the number of stable sites that can anchor Pt²⁺ species. The presence of metallic Pt can therefore be identified by both Pt 4f and C 1s spectra. The appearance of metallic Pt features in the Pt 4f spectra is concomitant with the appearance of a C 1s feature (F) corresponding to CO in the on-top configuration.

4. Reactivity of supported single atom catalysts towards H₂

4.1. Supported single metal atoms in the absence of metal particles

Dissociation of molecular hydrogen is the major function of the anode catalyst in hydrogen powered PEMFCs. It was reported that surface cationic Pt species promote dissociation of H₂ and facilitate hydrogen spillover on Pt-doped CeO₂ powders prepared by the solution combustion method.¹¹⁴ The origins of this reactivity, however, remained obscure.

One of the reasons is that the role of atomically dispersed Pt²⁺ on CeO₂ had not been studied individually, *i.e.* in the absence of other active species such as Pt⁴⁺, Pt⁰, and oxygen vacancies.^{114–116} Using well-defined model catalysts, atomically dispersed Pt²⁺ species can be prepared and studied individually.^{21,99} The corresponding approach involves the preparation of Pt–CeO₂ films with low Pt loading at low temperature followed by brief annealing to 700 K.²¹ The same procedure allows the preparation of atomically dispersed Pd²⁺ and Ni²⁺ species on nanostructured ceria.⁹⁸ Thus, the reactivity of atomically dispersed Pt²⁺, Pd²⁺, and Ni²⁺ species can be investigated under identical experimental conditions.

Changes in the oxidation states of Pt²⁺, Pd²⁺, and Ni²⁺ species upon reaction with H₂ were investigated by means of SRPES with high surface sensitivity under UHV conditions (see Fig. 19). Additionally, the oxidation state of Ce cations was monitored by means of resonant photoemission spectroscopy (RPES). The resonant enhancement ratio (RER) scales with the Ce³⁺/Ce⁴⁺ concentration ratio.¹¹⁷

Based on the analysis of the corresponding Pt 4f, Pd 3d, and Ni 2p core level spectra and the RER in Fig. 19, it was concluded that isolated Pt²⁺, Pd²⁺, and Ni²⁺ species do not facilitate dissociation of molecular hydrogen under the



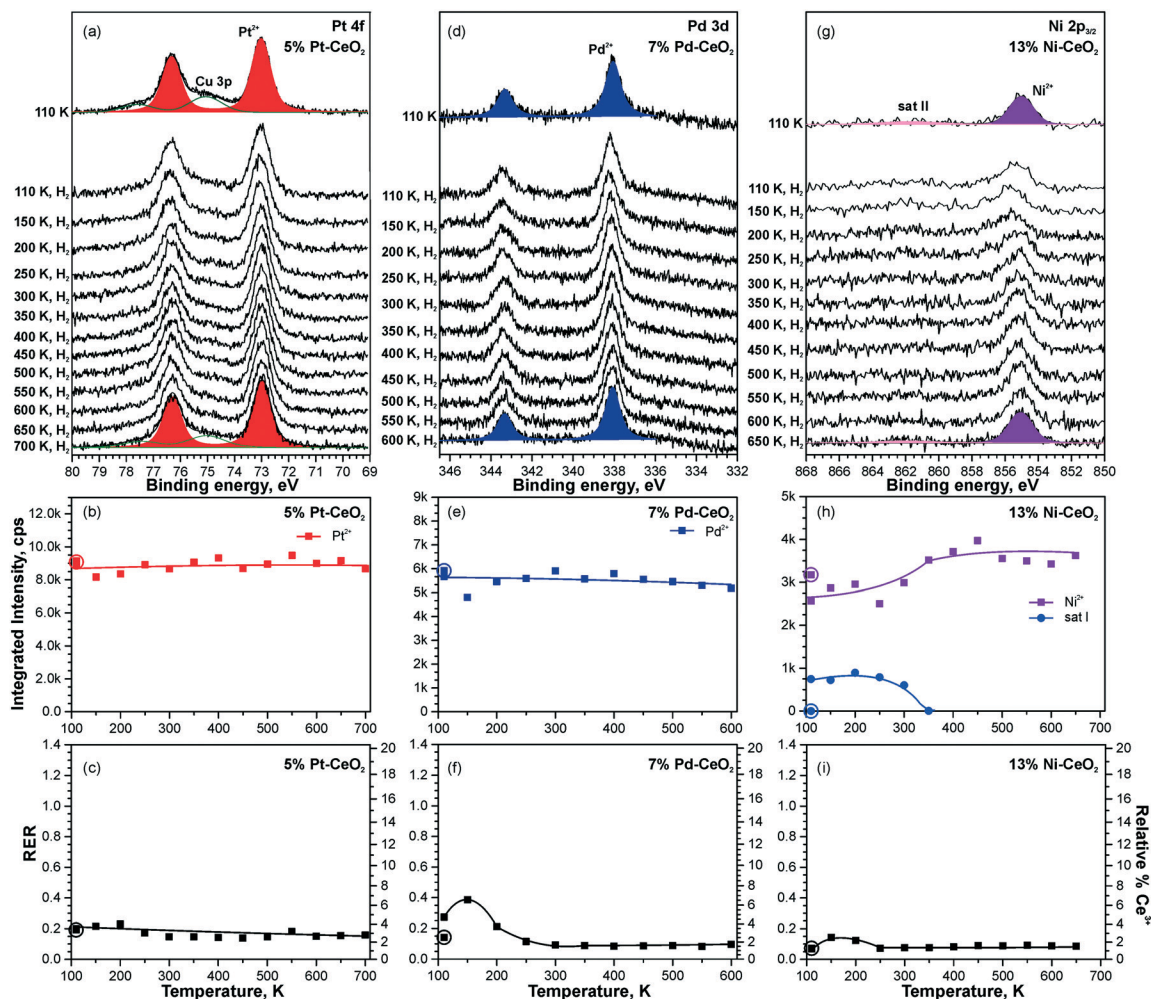


Fig. 19 Reactivity of 5% Pt–CeO₂ (a–c), 7% Pd–CeO₂ (d–f), and 13% Ni–CeO₂ (g–i) films towards hydrogen activation. The samples were briefly annealed at 700 K (5% Pt–CeO₂, 13% Ni–CeO₂) or 600 K (7% Pd–CeO₂). Pt 4f (a), Pd 3d (d), and Ni 2p_{3/2} (g) spectra obtained with $h\nu = 180$ eV, $h\nu = 410$ eV, and $h\nu = 1000$ eV, respectively. The integrated intensities of the surface species (b, e and h) and RER (c, f and i) on 5% Pt–CeO₂ (b and c), 7% Pd–CeO₂ (e and f), and 13% Ni–CeO₂ (h and i) following the exposure to 50 L of H₂ as a function of temperature. The data points obtained prior to the hydrogen exposure are circled. (a–c) Reproduced from ref. 99 with permission from the PCCP Owner Societies. (d–i) Adapted with permission from ref. 98, Copyright 2016, American Chemical Society.

experimental conditions employed, since reduction of neither these metal cations nor Ce⁴⁺ ones by H₂ has been detected.^{98,99}

This experimental finding was corroborated by DFT calculations for the case of ceria supported Pt²⁺ species.⁹⁹ The reaction pathways and the energetics for H₂ dissociation on the Pt–CeO₂ system and the CeO₂(111) surface are compared in Fig. 20.

The Pt–CeO₂ system was modeled as a low-energy vicinal surface exposing CeO₂(111) terraces and step edges along the [110] direction accommodating Pt²⁺ ions. These calculations showed that the barrier for H₂ dissociation at the Pt²⁺ sites (1.2 eV) is even larger than the barrier for dissociation on the pristine CeO₂(111) surface (1.0 eV).⁹⁹ The rather high H₂ activation energy is consistent with a very high thermodynamic stability of the Pt²⁺ species.²¹ This implies that the low barrier desorption of weakly adsorbed H₂ is strongly favored over dissociation, in accordance with the experimental findings.⁹⁹

In contrast to isolated Pt²⁺ species, DFT simulations show that sub-nanometer Pt clusters supported on ceria are highly active for H₂ dissociation. For instance, the calculated activation energy of only ~0.3 eV (Fig. 20, green line) for Pt₆ on CeO₂(111) could be easily overcome even at room temperature.⁹⁹

4.2. Supported single atom catalysts with traces of metallic clusters

Higher levels of metal loading in Pt–CeO₂ and Pd–CeO₂ films yield metallic particles upon annealing⁹⁸ (Fig. 21a and d). In contrast, no metallic Ni was detected in Ni–CeO₂ films despite the high metal loading (Fig. 21g).

The activation of molecular hydrogen becomes strongly favored in the presence of metallic Pt and Pd particles on Pt–CeO₂ and Pd–CeO₂ films. The process is accompanied by the reduction of Pt²⁺ and Pd²⁺ species coupled with the reduction



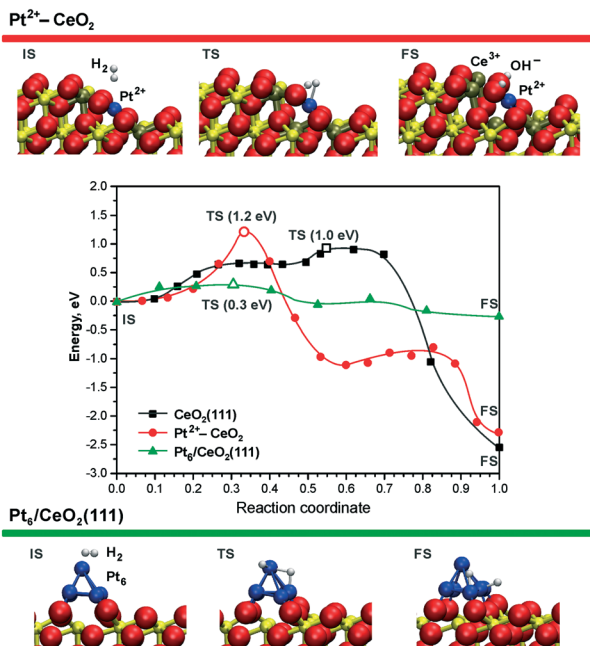


Fig. 20 Minimum energy paths for H₂ dissociation on the pristine CeO₂(111) surface (black), on the ionic Pt²⁺ sites (red), and on the supported Pt₆ cluster (green) – middle panel. The top and bottom panels display the initial (IS), transition (TS), and final (FS) states of H₂ dissociation on the Pt²⁺-CeO₂ and Pt₆/CeO₂(111) systems. The activation energies in TS states (open symbols) are given in parentheses. Ce⁴⁺, Ce³⁺, and O²⁻ ions are displayed as yellow, brown, and red spheres, respectively. Reproduced from ref. 99 with permission from the PCCP Owner Societies.

of Ce⁴⁺ at elevated temperatures (Fig. 21). Eventually, all Pt²⁺ and Pd²⁺ species are reduced quantitatively to metallic Pt⁰. Remarkably, very small amounts of metallic Pt (the traces hardly detected by conventional XPS) turned out to be sufficient to initiate H₂ dissociation under similar conditions.⁹⁹

In contrast to the Pt-CeO₂ and Pd-CeO₂ films, no reactivity towards H₂ was observed on the Ni-CeO₂ mixed oxide regardless of the Ni loading. The reason is the particularly high stability of the Ni²⁺ state preventing formation of metallic Ni.⁹⁸ In order to probe the role of metallic Ni in hydrogen activation, small amounts of metallic Ni were deposited onto Ni-CeO₂ films in UHV.⁹⁸ However, annealing of the film containing metallic Ni⁰ in hydrogen at 600 K did not yield a detectable change of the oxidation state of Ni. Apparently, the reaction sequence leading to reduction of cationic Pd and Pt species in Pd-CeO₂ and Pt-CeO₂ films is not possible for Ni-CeO₂, even in the presence of metallic Ni. Besides the high stability of Ni²⁺, a reason for the low reactivity could be that annealing of the Ni particles on Ni-CeO₂ triggers the formation of a NiO capping layer. Such encapsulation processes of the metallic Ni particles could prevent the dissociation of hydrogen.

Overall, the Pt-CeO₂ and Pd-CeO₂ films show a very similar reactivity towards H₂. Therefore, the mechanisms of hydrogen activation on both Pd-CeO₂ and Pt-CeO₂ films are expected to be similar. With respect to the reduction of atomically dispersed Pt-group species, density functional model-

ing suggests that hydroxylation of the Pt-O₄ moieties upon adsorption of atomic hydrogen²¹ leads to the reduction of Ce⁴⁺ cations and may result in destabilization and, possibly, reduction of Pt²⁺ species. In the presence of metallic Pt⁰ on Pt-CeO₂, atomic hydrogen could spillover from Pt particles onto the support, leading to hydroxylation of the Pt-O₄ moiety. In fact, the calculated reaction pathway on Pt₆/CeO₂ (Fig. 20) suggests the tendency of H to migrate to the cluster periphery and accumulate at the boundary in contact with the oxide support.⁹⁹ However, this pathway was ruled out based on the insignificant changes in the RER at temperatures where hydrogen reverse spillover processes are typically observed,¹¹⁸ *i.e.* between 190 and 260 K (see Fig. 21c and f). This suggests that hydrogen spillover and hydroxylation of the PtO₄ moiety are not the key steps in Pt²⁺ reduction during the reaction with H₂.

Thus, the mechanism of H₂ dissociation involving Pt²⁺ reduction is likely to be associated with the formation of oxygen vacancies upon reverse oxygen spillover¹¹⁹ from Pt-CeO₂ to the Pt particles. Upon exposure to H₂ the spilt-over oxygen is continuously removed from the Pt particles by reaction with hydrogen and subsequent formation of water. This reaction channel leads to the formation of oxygen vacancies accompanied by the reduction of Ce⁴⁺.

5. Redox conversion of atomically dispersed species to sub-nanometer particles

5.1. The role of oxygen vacancies

Density functional calculations show that the formation of oxygen vacancies can notably lower the adsorption energy of dispersed Pt on the ceria nanoparticles (Fig. 22). The energetic stability of the Pt²⁺ species depends on the proximity of the oxygen vacancies, their number, the Pt²⁺ concentration, and the distribution of Ce³⁺ ions.⁹⁹ Reduction of Pt²⁺ species is expected once the adsorption energy of atomic Pt falls below the bulk cohesive energy of Pt (-5.85 eV (ref. 120)).

The formation of a single oxygen vacancy outside the PtO₄ moiety (Fig. 22a) is not sufficient to initiate the reduction of a Pt²⁺ cation. In contrast, removal of one oxygen atom directly from the PtO₄ moiety lowers the adsorption energy Pt²⁺ species strongly and triggers the reduction, followed by release of Pt from the resulting Pt-O₃ moiety. However, the vacancy formation is strongly disfavored at the pocket sites in the presence of Pt²⁺ species, rendering such a process rather unlikely.⁹⁹ This suggests that oxygen vacancies are preferentially formed outside the Pt-O₄ moiety. If two such oxygen vacancies are created per Pt²⁺ ion, the Pt²⁺ adsorption energy falls well below the cohesive energy (Fig. 22b). Thus, density functional calculations indicate that the onset of Pt²⁺ reduction occurs when approximately two oxygen vacancies are created per Pt²⁺ site.

A rough estimation of the ratio between the concentration of oxygen vacancies and the Pt²⁺ cations at the onset of Pt²⁺





Fig. 21 Reactivity of 18% Pt–CeO₂ (a–c), 16% Pd–CeO₂ (d–f), and 42% Ni–CeO₂ (g–i) films towards hydrogen activation. The samples were briefly annealed at 700 K (18% Pt–CeO₂, 42% Ni–CeO₂) or 600 K (16% Pd–CeO₂). Pt 4f (a), Pd 3d (d), and Ni 2p_{3/2} (g) spectra obtained with $h\nu = 180$ eV, $h\nu = 410$ eV, and $h\nu = 1000$ eV, respectively. The integrated intensities of the surface species (b, e and h) and RER (c, f and i) on 18% Pt–CeO₂ (b and c), 16% Pd–CeO₂ (e and f), and 42% Ni–CeO₂ (h and i) following the exposure to 50 L of H₂ as a function of temperature. The data points obtained prior to the hydrogen exposure are circled. (a–c) Reproduced from ref. 99 with permission from the PCCP Owner Societies; (d–i) adapted with permission from ref. 98, Copyright 2016, American Chemical Society.

reduction could be obtained also from experimental data.^{99,121} For the Pt–CeO₂ film, the corresponding ratio for reduction of Pt²⁺ to Pt with hydrogen was determined to be about 1.5 O vacancies per Pt²⁺ cation.

The direct involvement of metallic Pt particles in the reduction of Pt²⁺ species can be ruled out by using methanol as a reducing agent.¹²¹ The decomposition of methanol on CeO₂ films involves the formation of formaldehyde, CO, and oxygen vacancies.^{122–124} Therefore, the reduction of Pt²⁺ species is observed as a consequence of the formation of oxygen vacancies regardless of the Pt loading (see Fig. 23). The estimation of the number of oxygen vacancies at the onset of Pt²⁺ reduction suggested that the formation of two oxygen vacancies is required for reduction of one Pt²⁺ species.¹²¹ This agrees with results of density functional calculations.⁹⁹ Also, a good correspondence of the ratio between the number of O vacancies and the number of reduced Pt²⁺ species determined in the

presence and in the absence of metallic Pt particles rules out a direct involvement of Pt particles in the destabilization of Pt²⁺.

Notably, the reduction of Pt²⁺ species on the Pt–CeO₂ film yields two types of metallic Pt, labeled as Pt⁰ and Pt* (see Fig. 23). The different binding energies are associated with differences in particle sizes. Based on the value of the binding energy in the Pt 4f spectra, the particle size for the Pt* components falls into the sub-nanometer range (less than 25 Pt atoms per particle) while the Pt⁰ signal corresponds to “regular” nanoparticles with a size above 2 nm.¹¹⁷

As one would expect, the formation of sub-nanometer Pt particles precedes the formation of larger Pt⁰ particles. The stability of the sub-nanometer Pt* particles strongly depends on the Pt loading in the Pt–CeO₂ films. In the presence of Pt⁰ particles, Pt* particles are rapidly lost by sintering and coalescence with Pt⁰ particles.





Fig. 22 Adsorption energies of atomic Pt in the 2+ state, E_{ad} , without ($n = 0$) and with n ($n > 0$) oxygen vacancies on a) Pt-Ce₄₀O_{80- n} and b) 4Pt-Ce₄₀O_{80- n} nanoparticle models. E_{ad} values calculated in the absence of oxygen vacancies are shown in bold. The transparent planes cut the nanoparticles into two halves. The vacancies with the signs “+” and “-” are located in front of and behind the planes, respectively. The sign “±” indicates two symmetric oxygen vacancies located in front of and behind the transparent planes. For the 4Pt-Ce₄₀O_{80- n} nanoparticle (b), the formation of one oxygen vacancy at the position +2 leads to a significant decrease of E_{ad} of one Pt²⁺ in the proximity of the vacancy (1Pt) while the remaining three Pt²⁺ species (3Pt) are influenced less significantly. Formation energies of oxygen vacancies, E_{vac} , in the Pt-Ce₄₀O₈₀ model (a) were calculated with respect to $\frac{1}{2}$ of the energy of the O₂ molecule. Reproduced from ref. 99 with permission from the PCCP Owner Societies.

In the absence of Pt⁰ particles, however, the sub-nanometer Pt* particles show remarkable stability upon annealing even under strongly reducing conditions (see Fig. 23a).

Due to the high binding energy of the Pt* contribution in the Pt 4f spectra (72.0 eV), the metallic nature of sub-nanometer Pt* particles is difficult to prove.^{98,99} For example, the presence of a PtO phase or the adsorption of CO on Pt gives rise to spectral contributions at similar binding energies.^{112,125} Therefore, it is instructive to compare the trends in the formation of oxygen vacancies on Pt-CeO₂ films with different Pt loadings with respect to the Pt-free CeO₂ film during reaction with methanol (see Fig. 24). The nearly linear increase of the RER on the Pt-free CeO₂ film indicates the formation of oxygen vacancies resulting from desorption of formaldehyde, CO, and water that involves partial removal of lattice oxygen from ceria.¹²²⁻¹²⁴ In the presence of metallic Pt particles, however, a second channel opens that involves reverse hydrogen spillover followed by desorption of molecular hydrogen.¹¹⁸

This process is accompanied by re-oxidation of ceria. As a result, the amount of the oxygen vacancies formed in the

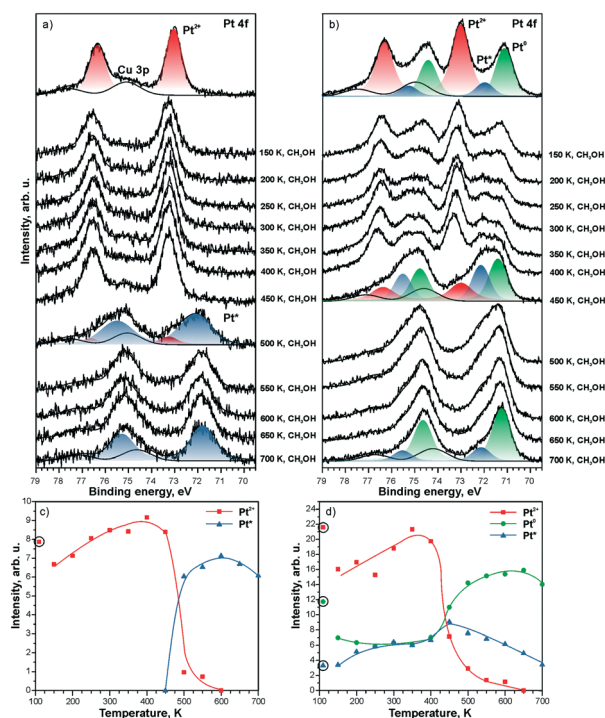


Fig. 23 Pt 4f spectra obtained from a) 5% Pt-CeO₂ and b) 18% Pt-CeO₂ films (top spectra) after annealing while exposing to methanol (10 L) between 150 and 700 K. The Cu 3p signal originates from the underlying Cu(111) substrate. The spectra were acquired with a photon energy of 180 eV. Integrated intensities of Pt²⁺, Pt⁰, and Pt* species on c) 5% Pt-CeO₂ and d) 18% Pt-CeO₂ films are shown as a function of temperature during annealing under methanol exposure. The circled data points were obtained prior to the methanol exposure. Reprinted from ref. 121, Copyright 2016, with permission from Elsevier.

presence of Pt particles is lowered. The limited density of oxygen vacancies in the presence of the Pt particles is reflected by the RER which shows saturation with increasing temperature. Notably, the formation of vacancies is limited on both Pt-CeO₂ films regardless of the Pt loadings. This suggests the

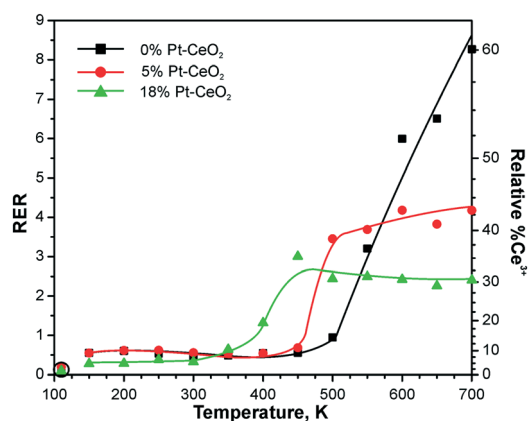


Fig. 24 RERs on 0% Pt-CeO₂ (black), 5% Pt-CeO₂ (red), and 18% Pt-CeO₂ (green) films as a function of the temperature during annealing under exposure to methanol (10 L). The data points indicated by circles were obtained prior to the methanol exposure. Reprinted from ref. 121, Copyright 2016, with permission from Elsevier.



presence of metallic Pt on all Pt–CeO₂ films and corroborates the assignment of the Pt* component to sub-nanometer Pt particles. It can be concluded that the formation of stable sub-nanometer Pt particles is possible *via* preparation of stable atomically dispersed Pt²⁺ species followed by their reduction. A similar preparation strategy was employed to obtain finely dispersed Pt particles *via* Pt segregation from Pt-doped ceria.¹²⁶

5.2. Redox coupling with the reducing agent

The relationship between the formation of oxygen vacancies and stability of the Pt²⁺ species^{99,121} enables new methods to prepare thermally stable sub-nanometer Pt particles.¹²⁷ Since the formation of oxygen vacancies in ceria-based materials leads to reduction of Ce⁴⁺ to Ce³⁺,^{123,128,129} one could speculate that reduction of Pt²⁺ could be controlled *via* the formation of Ce³⁺ centers.

Following this idea, the reduction of Pt²⁺ species in Pt–CeO₂ materials might be triggered by addition of reducing agents

which do not involve the removal of oxygen. For instance, the adsorption of Sn on CeO₂ films results in the formation of Sn²⁺ cations accompanied by the reduction of Ce⁴⁺ cations and formation of Ce³⁺ centers.^{130–132} Thus, the adsorption of Sn on Pt–CeO₂ could trigger the conversion of Pt²⁺ to sub-nanometer Pt particles. The process could occur at low temperature and without the formation of oxygen vacancies.

Indeed, stepwise Sn deposition onto the Pt–CeO₂ film at 300 K in UHV initiated the reduction of Pt²⁺ species yielding sub-nanometer Pt particles (see Fig. 25). Comparing the trends in the formation of Pt particles under different Sn deposition conditions, it was possible to correlate the onset of Pt²⁺ reduction with the concentration of the Ce³⁺ centers determined by the RER.¹²⁷ Estimating the number of Ce³⁺ centers formed by Sn deposition on the Pt–CeO₂ film it was concluded that approximately 6 Ce³⁺ cations are required per Pt²⁺ species to induce the reduction of Pt²⁺. This number corresponds to 3 Sn²⁺ cations deposited per Pt²⁺ species. This finding is in good agreement with density functional modeling data.¹²⁷

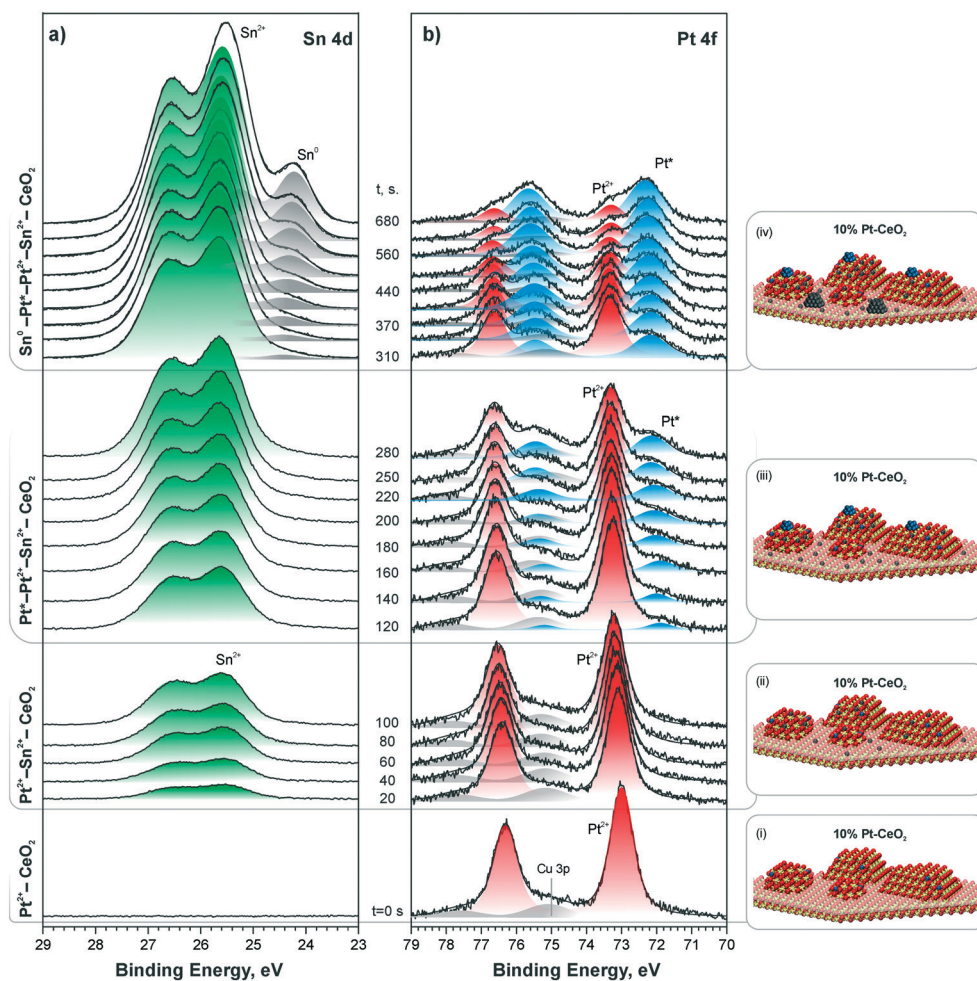


Fig. 25 Sn 4d (a) and Pt 4f (b) spectra obtained from a 10% Pt–CeO₂ film during stepwise Sn deposition at 300 K in UHV. The Sn 4d and Pt 4f spectra were obtained with photon energies of 60 and 180 eV, respectively. The spectra are grouped (i)–(iv) according to the Pt and Sn oxidation states in the film. In the corresponding ball-and-stick models (i)–(iv), red, yellow, blue, and dark grey balls represent oxygen, cerium, platinum, and tin, respectively. Reproduced from ref. 127 with permission from the Royal Society of Chemistry.



6. Properties of supported sub-nanometer metal particles

6.1. Metal–support interaction as a function of particle size

Metal particles supported on reducible ceria-based materials retain a number of unique properties arising from electronic metal–support interactions (EMSI).^{34,102,119,133} These include, for instance, higher reactivity and selectivity, structural flexibility, and self-cleaning capacity. Depending on the strength of the EMSI, there is a substantial charge transfer between the supported metal particle and the supporting oxide.^{73,117,119,134,135} Combining synchrotron-radiation photoelectron spectroscopy and scanning tunneling microscopy it was possible to “count” the number of electrons transferred across the metal/oxide interface as a function of particle size (Fig. 26).¹¹⁷ The maximum charge transferred per Pt atom was detected at particle sizes between 1 and 1.5 nm corresponding to a number of 30 and 70 Pt atoms per particle (see Fig. 26a). At higher Pt coverage the overall number of transferred electrons approaches a limit (Fig. 26b). Consequently, the average charge per Pt atom decreases as a function of size for particles containing more than 70 Pt atoms (Fig. 26a). At these particle sizes, the metal–support interaction is mostly

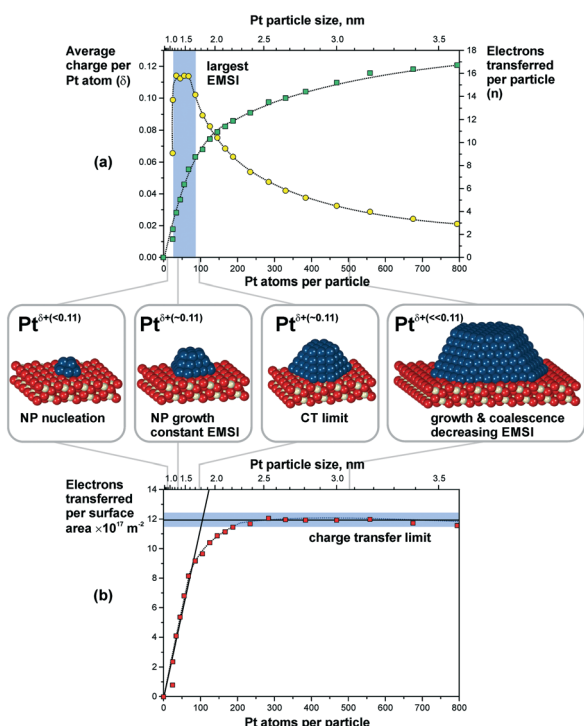


Fig. 26 (a) The number of electrons transferred per Pt particle to the ceria support increases with increasing particle size (green squares). The partial charge per Pt atom reaches a maximum for particles with 30 to 70 atoms (yellow circles). Here, the EMSI is largest. (b) At higher Pt coverage the total amount of transferred charge approaches a limit which we denote as the “charge transfer limit” (red squares). The atomic models show schematically the average particle sizes in the different regions. Reproduced with permission from ref. 117, Copyright 2016, Nature Publishing Group.

associated with the restructuring of the particle shape and Pt/ceria interface which is rather insensitive to the charge transfer.¹³⁶ Density functional modeling rationalized the magnitude of the charge transfer, mainly depending on three parameters, *i.e.* the size of the Pt particles, their density, and the presence of oxygen vacancies¹¹⁷ (Fig. 27). In particular, the limits of charge transfer were attributed to the electrostatic destabilization at a high surface concentration of Ce^{3+} centers. In this respect, the degree of nanostructuring, *i.e.* the size of ceria nanoparticles or the roughness of the ceria support, becomes an additional parameter that has a direct impact on the magnitude of the charge transfer.¹³⁴ The charge transfer would thus increase with increasing density of low-coordinated Ce^{4+} cations which are capable of accepting electrons from Pt particles and still avoid electrostatic destabilization.¹³⁴

The most remarkable behavior was observed in the regime of sub-nanometer sized Pt particles (Fig. 26a). There, the charge transfer per Pt atom was found to decrease compared to what was predicted by DFT calculations for Pt particles on surfaces of stoichiometric ceria.^{117,119,137} The observed phenomenon was assigned to nucleation of Pt at Ce^{3+} defect sites.^{117,138} Indeed, calculated models also showed that vacancies reduce the net charge transfer from Pt to ceria (Fig. 27). In consequence, the net charge transfer is a function of the degree of reduction of the ceria support.^{35,139,140} Depending on the specific system, metal particles can bear a positive charge on highly oxidized ceria supported and a negative charge on highly reduced supports.^{139,140}

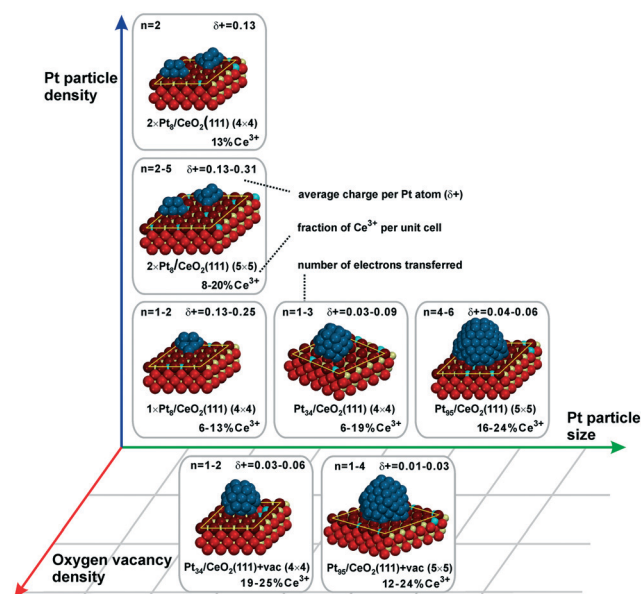


Fig. 27 Summarized results of density functional calculations of $\text{Pt}_n/\text{CeO}_2(111)$ models. The particle size, the particle density, and the oxygen vacancy density on the support are important factors that control the charge transfer across the metal–oxide interface. Reproduced with permission from ref. 117, Copyright 2016, Nature Publishing Group.



In summary, the charge transfer is controlled by the particle size, the particle density, the support structure and the degree of reduction of the ceria support.

The DFT modeling suggests that the charge is mainly localized on metal atoms in the vicinity of the interface.^{117,136} Small Pt aggregates in the sub-nanometer regime consist of only a few atomic layers and, therefore, most of their surface atoms bear an excess charge. This effect may be employed to modify the catalytic properties of the supported particles.

6.2. Charge transfer in aqueous environments

On ceria-based catalysts, water molecules readily dissociate *via* proton transfer to surface O atoms, resulting in partial surface hydroxylation.^{101,141,142} DFT calculations were employed to investigate water dissociation at the ceria (111) surface¹⁴³ and at the ceria/liquid interface¹⁴⁴ as well as to elucidate the charge transfers between Pt particles and the ceria support resulting from water dissociation. The influence of the aqueous environment on the reaction mechanism, thermodynamics and kinetics was investigated by means of *ab initio* molecular dynamics (MD) simulations.¹⁴⁴

At equilibrium, these MD simulations revealed the existence of a fast proton dynamics at the water/ceria interface involving surface hydroxyl groups, solvated hydroxide ions, and water molecules coordinated to the surface Ce⁴⁺ cations at the water/ceria interface.¹⁴⁴ Under these conditions, the dissociation of water molecules in the first surface solvation layer becomes a reversible dynamic process, which is governed by solvent-induced short-ranged transfers of protons between the adsorbed water molecules and the surface oxygen sites, or, in the reverse process, between surface hydroxyl groups and solvated hydroxide ions. The local increase of concentration of hydroxide ions on ceria surface, which results from water dissociation, activates a water-mediated Grotthus-like mechanism that gives rise to fast long-ranged proton diffusion along the water/ceria interface.¹⁴⁴ It involves concerted transfer of protons along the chains formed by solvent water molecules bridging between surface hydroxide and surface water molecules (Fig. 28a and b).

The catalytic importance of this dynamic process was demonstrated for water dissociation in the presence of supported sub-nanometer-sized Pt nanoparticles where the solvent accelerates the spillover of ad-species between oxide and metal sites.¹⁴⁴ The mechanism of this process, calculated in the presence of a Pt₆ cluster on a stoichiometric ceria (111) slab, is demonstrated in Fig. 28c–e. In the first step, a water molecule of the solvation layer dissociates at the periphery of the supported Pt particle into a hydroxide ion and a proton. This activates the dissociation of a neighboring solvent water molecule (see ovals in Fig. 28c and d), which mediates the proton transfer at the water/oxide interface. The proton is effectively transferred to an oxide surface site away from the nanoparticle, forming a surface hydroxyl (Fig. 28e). The resulting OH[−] species readily binds to a Pt site of the

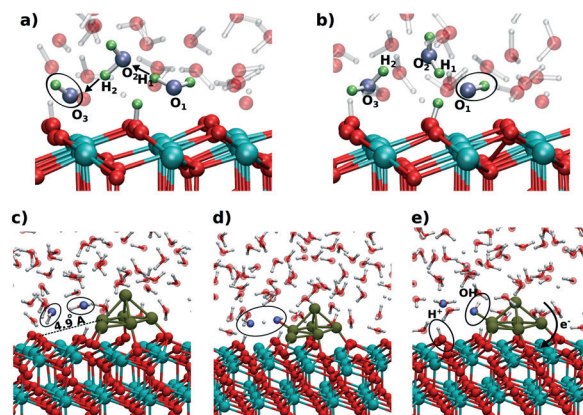


Fig. 28 Snapshots from the *ab initio* MD simulations showing the initial (a and c), intermediate (d), and final (b and e) configurations of the proton chain leading to long-range proton diffusion on the CeO₂(111) surface (a and b) and at the periphery of a supported Pt₆ cluster (c–e). Adapted with permission from ref. 144, Copyright 2016, American Chemical Society.

supported cluster inducing substantial charge transfer across the metal/oxide interface.

Upon electron reorganization in the Pt/CeO₂ system one electron from the hydroxide species and another one from the Pt cluster are transferred to the oxide substrate through the metal/oxide interface, thus indicating that the cluster/solvent and metal/oxide interfaces are strongly coupled.¹⁴⁴ This demonstrates that, with respect to UHV conditions, the number of electrons transferred from a supported Pt cluster in an aqueous environment to the ceria support is larger by up to a factor of two.¹⁴⁴ Given the concerted reaction mechanism, the transfer of protons away from the Pt cluster cannot be distinguished from the transfer of the hydroxide ion towards the Pt cluster. In any case, however, the supported Pt clusters act as a basin attracting hydroxide species, which consequently accumulate at the cluster sites.

6.3. Formation of oxygen vacancies and reverse oxygen spillover

In the context of metal–support interaction, the supported metal particles may have a significant influence on the reactivity of ceria and, particularly, on its oxygen storage capacity. For bare ceria, the formation of oxygen vacancies in ceria nanostructures is strongly favored with respect to bulk samples.^{65,68,83,145,146} DFT calculations suggest that the oxygen vacancy formation energy decreases from 2.25 eV on the CeO₂(111) slab to 0.8 eV on ceria nanoparticles such as Ce₄₀O₈₀ (see Fig. 29). The deposition of a Pt particle appears to have only little effect on the energy of vacancy formation, both on extended ceria^{73,119} and on ceria nanoparticles¹³⁴ (Fig. 29). However, the formation of an oxygen vacancy results in a variation of the initial charge of the supported Pt₈ particle as a function of the size of the ceria nanoparticle, *i.e.* the degree of ceria nanostructuring.^{119,134} This effect reflects the flexibility of the electronic structure of nanostructured ceria-based systems.



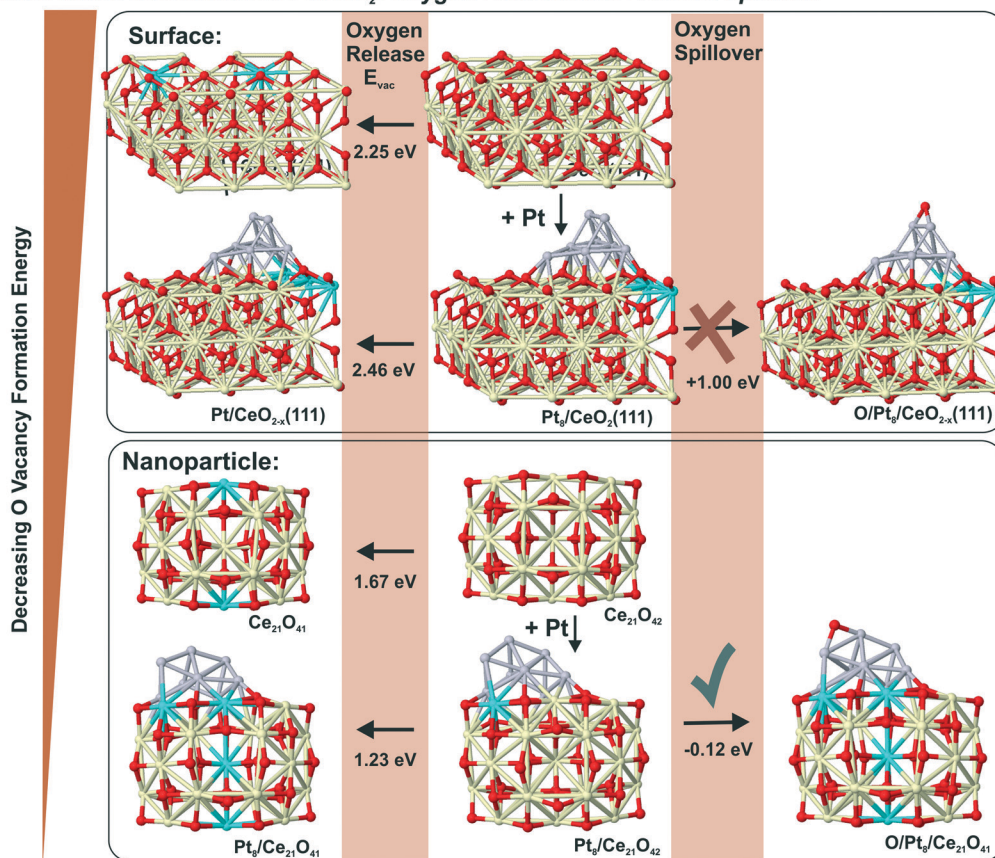
Metal Oxide Interaction on Pt/CeO₂: Oxygen Release and Reverse Spillover

Fig. 29 Oxygen release and oxygen reverse spillover in extended and nanostructured Pt/ceria models according to DFT calculations. Reproduced with permission from ref. 119, Copyright 2011, Nature Publishing Group.

In the presence of supported Pt particles, oxygen vacancies can also be created by migration of atomic oxygen from the support onto Pt particles. The corresponding process is denoted as reverse oxygen spillover. Considering the energy of oxygen vacancy formation on CeO₂(111) (2.25 eV) and the adsorption energy of oxygen on Pt(111) (0.60–1.55 eV),¹⁴⁷ such a reverse spillover process seems to be energetically unfavorable. At this point, the degree of the nanostructuring of ceria plays a critical role. In particular, the low energy for oxygen vacancy formation on nanostructured ceria favors reverse oxygen spillover and makes the process exothermic by 0.51 eV (Fig. 29). The reaction is accompanied by the formation of two new Ce³⁺ centers. These estimations strongly suggest that the occurrence of oxygen reverse spillover intrinsically requires the presence of a nanostructured ceria support.

The above described processes associated with charge transfer and reverse oxygen spillover were investigated experimentally (see Fig. 30). In particular, the initial increase in Ce³⁺ concentration indicates the charge transfer from the Pt particles to the ceria film. In contrast to the charge transfer, reverse oxygen spillover is associated with a substantial activation barrier and does not occur spontaneously at low temperature. However, a rapid increase of the Ce³⁺ concentration was observed upon annealing above 500 K which is associated with the onset of reverse oxygen spillover (Fig. 30).

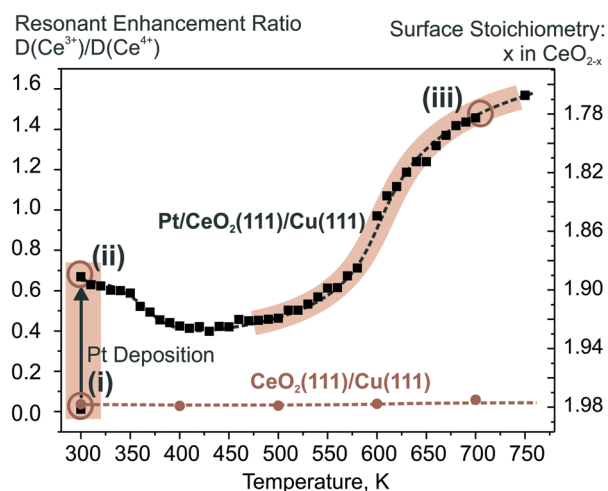


Fig. 30 The evolution of the resonant enhancement ratio (RER) indicates the occurrence of the charge transfer upon Pt deposition on nanostructured CeO₂(111) at 300 K and the reverse oxygen spillover above 500 K. The RER determined from the heights of the resonant features in the valence bands associated with Ce³⁺ and Ce⁴⁺ cations scales with the Ce³⁺/Ce⁴⁺ concentration ratio by a factor of 5.5.¹¹⁷ Reproduced with permission from ref. 119, Copyright 2011, Nature Publishing Group.



6.4. Structural flexibility of sub-nanometer metal particles

The structural flexibility of metal particles facilitates a chemical reaction by reducing the activation barrier.^{137,148–153} In the regime of sub-nanometer size this property should become increasingly important. For instance, for a Pt₆ cluster supported on a regular CeO₂(111) surface the structural flexibility of the former turns the energetic balance of oxygen reverse spillover from endothermic, when structural changes are neglected, to weakly exothermic, when structural changes are taken into account (see Fig. 31).¹³⁷ In other words, the binding energy of an oxygen atom to the Pt₆ cluster becomes comparable to the binding energy of oxygen in the ceria lattice if structural rearrangements in the Pt cluster are considered. Interestingly, this effect suggests that oxygen transfer to supported Pt clusters does not necessarily require nanostructured ceria particles, but may become slightly exothermic for very small Pt clusters even on extended ceria surfaces.^{137,145} Notably, the initial and final stages of this process identified with global structural optimization indicate a transition from a 3D to 2D morphology of the supported Pt₆ cluster (Fig. 31).

6.5. Interconversion between atomically dispersed Pt²⁺ and sub-nanometer sized Pt particles

Because of the strong correlation between the oxidation state of Pt and the degree of reduction of the ceria support, it is expected that the Pt–CeO₂ catalyst can switch between two stable states, one associated with atomically dispersed Pt²⁺ species anchored at {100} sites and a second one with the Pt

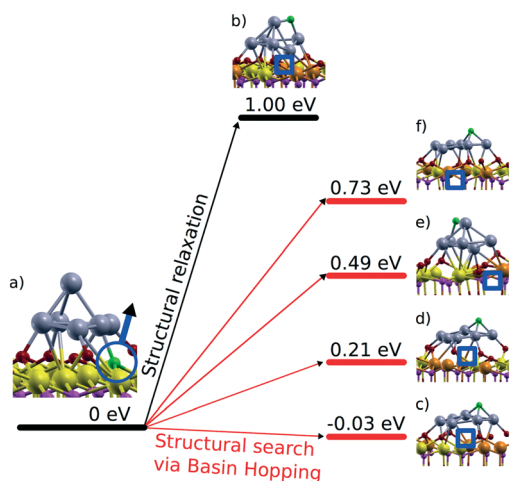


Fig. 31 Thermodynamics of reverse oxygen spillover between a supported Pt₆ cluster and a CeO₂(111) surface. Initial state (a) Pt₆ cluster deposited on a stoichiometric CeO₂(111) slab and final states (b–f), partially oxidized Pt₆O cluster on the slab with an oxygen vacancy beneath the cluster, resulting from a usual structural relaxation towards the local minimum (b) and from a basin hopping structure optimization (c–f) that allows the global minimum to be reached in the configurational space. Pt atoms are in gray, oxygen from the first (second) layer in red (violet), Ce⁴⁺ in yellow, and Ce³⁺ in orange. The oxygen vacancy site is indicated by blue squares. Reproduced with permission from ref. 137, Copyright 2014, American Chemical Society.

from sub-nanometer-sized Pt aggregates. The interconversion between the two states will occur in response to changes of the reactive environment. The latter may be associated with changes in the gas composition or, in the case of the fuel cell catalysts, in response to changes of the electrode potential.

Reversible changes in the oxidation state of Pt particles supported on nanostructured ceria have previously been monitored by Hatanaka *et al.*¹⁵⁴ The presence of oxidized Pt species was explained by the formation of Pt–O–Ce bonds in an oxidative atmosphere at high temperature. It was proposed that the formation of such bonds drives the re-dispersion of Pt particles into smaller clusters.¹⁵⁵ These studies^{154,155} suggested the formation of a monolayer of oxidized Pt on the surface of a ceria-based material after treatment in an oxygen atmosphere. Based on the binding energy of the corresponding species in Pt 4f spectra, we speculate that the species described in that work are similar to Pt²⁺ species anchored at {100} facets described above.

On single-crystal-based Pt–CeO₂ model films containing traces of metallic Pt, annealing in a H₂ atmosphere was shown to trigger reduction of Pt²⁺ and formation of metallic Pt, both in the form of sub-nanometer sized aggregates and larger Pt particles.^{98,99} The corresponding situation is shown in Fig. 32a.

Remarkably, the subsequent annealing of the Pt–CeO₂ film in oxygen triggers partial re-dispersion of metallic Pt particles. We observe a partial recovery of the characteristic signal associated with the Pt²⁺ species anchored at the {100} sites (Fig. 32a). Similar and partially reversible changes were also observed between Pd²⁺ species and metallic Pd particles (Fig. 32b). The degree of reversibility between two oxidation states likely depends on the conditions of the redox treatment, *e.g.* the temperature and the pressure of hydrogen and oxygen. Additionally, the degree of sintering can be controlled by the metal loading in a ceria-based catalyst.¹⁵⁶ In this respect the ideal Pt loading should allow full recovery of the atomically dispersed state.¹⁰⁴ It is likely that the degree of stabilization and the strong dependence of the catalytic activity (power density) on Pt loading observed for the Pt–CeO₂

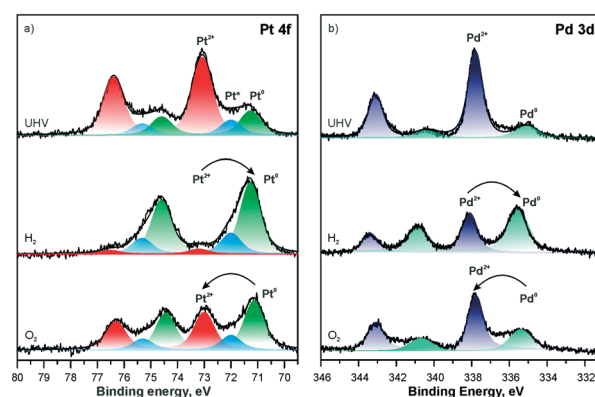


Fig. 32 The development of Pt 4f (a) and Pd 3d (b) spectra obtained from Pt–CeO₂ and Pd–CeO₂ films, respectively, following the annealing in UHV (top), under H₂ (middle), and O₂ (bottom) atmosphere.



anode fuel cell catalysts are associated with the degree of reversibility that can be achieved in the catalyst film.⁴¹

7. Active state of Pt–CeO₂ thin film catalysts under electrochemical conditions

The active state of the Pt–CeO₂ catalyst films under electrochemical conditions can be examined by means of spectroelectrochemical methods. In line with the above discussion, changes of the electrode potential give rise to dynamic changes in the oxidation state of ceria and, in response, also of the supported Pt species. The stability of Pt–CeO₂ catalyst films was investigated as a function of Pt loading by means of electrochemical infrared spectroscopy.¹⁰⁹ In these experiments CO formed during electro-oxidation of methanol was employed as a probe molecule. The corresponding IR spectra obtained from Pt–CeO_x films are presented in Fig. 33. Reference spectra obtained from Pt(111) under identical conditions are also shown for comparison.

Typically, CO adsorption on Pt(111) gives rise to two bands associated with adsorption in on-top (CO_t) and bridging (CO_b) configurations (see Fig. 33a). Clearly, the emergence of CO_t stretching bands on Pt–CeO_x films indicates that a fraction of the atomically dispersed Pt²⁺ species was converted into metallic Pt particles already in the first potential cycle.

The red shift and broadening of the CO_t band along with the absence of the CO_b on Pt–CeO_x films suggests the formation of small Pt aggregates with less developed (111) facets.¹⁵⁷ In particular, the contributions from the low-coordinated sites and from (111) facet are well resolved on Pt–CeO_x films with high Pt loading (see dark and light grey areas and arrows in Fig. 33a). In the limit of low Pt loadings, the CO_t band is dominated by the contributions from CO adsorbed at low-coordinated sites. This situation is consistent with the formation of sub-nanometer-sized Pt particles that do not expose any measurable fraction of regular (111) facets.

Under electrochemical conditions, the blue shift of the CO_t bands with increasing electrode potential is associated with the Stark effect.¹⁵⁸ The relative offsets in the Stark plot between the Pt(111) and Pt–CeO_x films (Fig. 33b) are related to the size of supported Pt particles. Density functional modeling reinforces the correlation between the stretching frequency and the particle size, on both unsupported and supported Pt particles.^{109,110} In particular, the offset of the CO vibrational frequencies by 26 cm⁻¹ with respect to Pt(111) indicates the formation of sub-nanometer-sized particles containing about 30 Pt atoms per particle.

The findings both in UHV and under electrochemical conditions suggest that the conversion of atomically dispersed Pt²⁺ to sub-nanometer sized Pt particles transforms the Pt–CeO_x film into an electrocatalytically active state. Notably, the stability of the sub-nanometer Pt particles under electrochemical conditions is a function of the Pt loading. A progressive blue shift of the CO_t band on the Pt–CeO_x film with high Pt loading during repetitive cycling is associated with sintering and growth of larger Pt particles with defined (111) facets (Fig. 34a). In sharp contrast, such changes are not seen for Pt–CeO_x films with low Pt loading, even after repeated

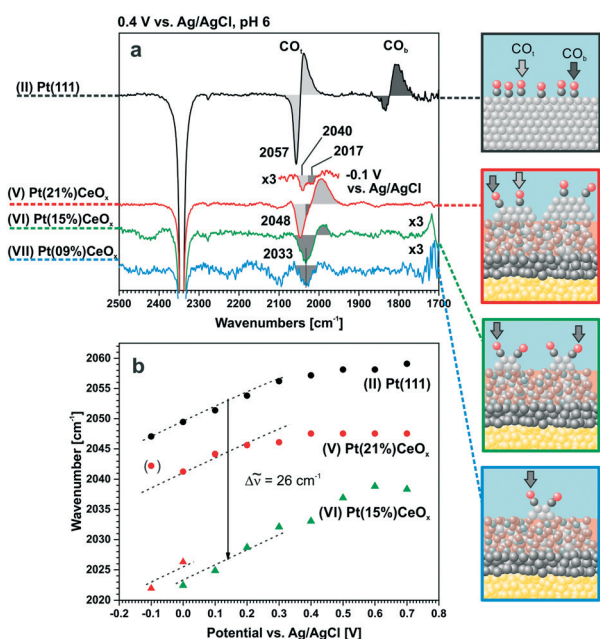


Fig. 33 Comparison of the electrochemical IR spectra taken during methanol oxidation on Pt–CeO_x electrocatalysts with different Pt concentrations at pH 6 (1 M CH₃OH in 0.1 M HClO₄): (a) IR spectra at an electrode potential of 0.4 V_{Ag/AgCl} (reference –0.15 V_{Ag/AgCl}); (b) peak position of the on-top CO band as a function of the electrode potential. The samples are numbered (I–VII) for easy comparison. Reproduced with permission from ref. 109, Copyright 2016, American Chemical Society.

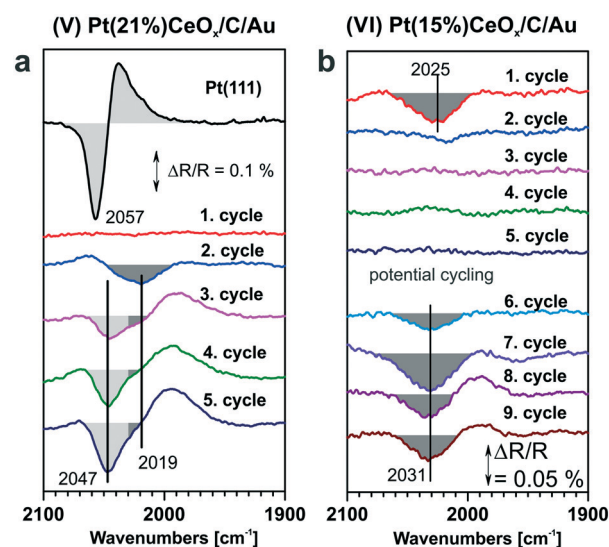


Fig. 34 Electrochemical IR spectra during repeated cycles of methanol oxidation on (a) Pt(21%)–CeO_x/C/Au and (b) Pt(15%)–CeO_x/C/Au at 0.4 V_{Ag/AgCl}. All reference spectra were taken at a potential of –0.15 V_{Ag/AgCl}. Adapted with permission from ref. 109, Copyright 2016, American Chemical Society.

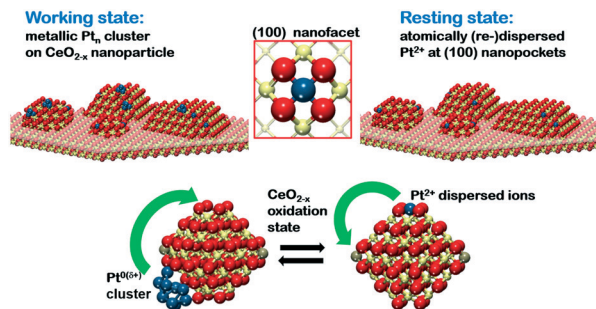


Fig. 35 The Pt–CeO_x catalyst in two stable states associated with atomically dispersed Pt²⁺ species (resting state) and sub-nanometer Pt particles (working state) in an electrochemical environment. The conversion between two states is driven by the charge transfer as a function of the electrode potential. DFT modeling corroborates reversible conversion as a function of the degree of reduction of ceria in the presence of {100} sites.

potential cycling (Fig. 34b). This behavior is consistent with enhanced stability of the sub-nanometer Pt particles on Pt–CeO_x films with low Pt loading in an electrochemical environment.

The enhanced stability of the sub-nanometer Pt particles is associated with the strong interaction between Pt particles and ceria support. The above findings suggest that Pt²⁺ species are formed under oxidizing conditions, which prevents the Pt particles from sintering during potential cycling. In this respect, the abundance of the {100} sites determines the maximum Pt loading in Pt–CeO_x films up to which the stabilization mechanism will operate. At Pt loadings exceeding the density of these sites, the formation and growth of larger metallic Pt particles is no longer suppressed.

In summary, the Pt–CeO_x catalyst with ideal Pt loading will be able to undergo reversible conversion between atomically dispersed Pt²⁺ species and sub-nanometer Pt particles as a function of the oxidation state and/or the electrode potential.

A consistent scenario would involve the presence of two stable states, denoted as the working state and the resting state (see Fig. 35). Accordingly, the resting state is associated with the formation of atomically dispersed Pt²⁺ anchored at {100} nanofacets. The conversion from the resting state into the working state is accompanied by the formation of sub-nanometer Pt particles triggered by the reduction under electrochemical conditions. Conversely, the resting state is recovered upon returning to oxidizing conditions. Note that such dynamic changes of the Pt oxidation state are typically associated with dissolution processes and sintering. In the presence of the ceria support, the release of Pt species may be prevented, leading to enhanced stability and constant activity.

Conclusions

A dynamic catalyst concept involving a single atom catalyst state has been applied to prepare anode fuel cell catalysts with high noble metal efficiency and enhanced stability. The approach involves the deposition of thin Pt–CeO₂ films on

carbon substrates by means of radio frequency magnetron sputtering from separate CeO₂ and Pt targets in a non-reactive atmosphere. The deposition process is accompanied by the etching of carbon substrates yielding highly porous columnar structures coated with Pt–CeO₂ films. Atomically dispersed Pt was found to be present in these films in the form of Pt²⁺ and Pt⁴⁺ species at the surface and in the bulk, respectively. The ratio between the Pt²⁺ and Pt⁴⁺ species strongly depends on the degree of the nanostructuring of the Pt–CeO₂ films. Fuel cell tests demonstrated the direct link between the electrocatalytic performance of the Pt–CeO₂ films and the density of the surface Pt²⁺ species.

The local structure and reactivity of the Pt²⁺ species supported on nanostructured ceria films have been investigated by a variety of techniques including high-resolution TEM, SRPES, AR XPS, HAXPS, and STM in combination with density functional modeling. The appropriate model systems were developed to identify possible anchoring sites for the atomically dispersed Pt species. By means of density functional calculations and surface-science-based model catalysts the adsorption sites yielding strongly anchored Pt²⁺ were associated with {100} nanofacets on ceria particles. There, the Pt²⁺ species were found to be located in the center of a planar PtO₄ moiety forming a quadratic coordination environment. A similar structural element was found at the steps of CeO₂(111) films. CeO₂(100) nanofacets were also identified on real Pt–CeO₂ catalyst films by means of HRTEM.

The stability and the reactivity of the Pt²⁺ species were investigated as a function of Pt loading on the model Pt–CeO₂ films under ultra-high vacuum conditions. It was found that the capacity of the nanostructured ceria surface to form isolated Pt²⁺ is limited by the number of available {100} sites. If the Pt loading exceeds the density of anchoring sites available, the growth of Pt nanoparticles is observed.

The activation of molecular hydrogen at single Pt²⁺ species is associated with a high activation barrier. In the presence of metallic Pt, the hydrogen activation is facile and eventually leads to reduction of all Pt²⁺ species. It was found that metallic Pt is not directly involved in the reduction of the Pt²⁺ species. Instead the reduction of Pt²⁺ is coupled to the reduction of Ce⁴⁺ to Ce³⁺, resulting from the formation of oxygen vacancies. Reaction with methanol also leads to reduction of the ceria support and triggers the formation of sub-nanometer Pt particles. Alternatively, the reduction of Pt²⁺ and the formation of Pt particles can also be mediated by co-adsorption of reducing agents such as Sn that lead to formation of Ce³⁺ without formation of oxygen vacancies. In an electrochemical environment, conversion of the Pt²⁺ species to sub-nanometer Pt particles is triggered by the electrode potential.

The sub-nanometer Pt particles exhibit remarkable thermal stability with respect to sintering under both UHV and electrochemical conditions. The chemical properties of the supported sub-nanometer-sized particles are modified by metal–support interactions associated with strong charge transfer, by their structural flexibility, and by spillover phenomena. Regarding



the electronic metal support interactions between Pt particles and ceria, the amount of charge accumulated on the Pt particle depends on the Pt particle, the particle density, the structure and the stoichiometry of the ceria support.

From the above results, a consistent picture emerges with respect to the operation of the Pt–CeO₂ catalyst films under electrochemical conditions. The stabilization mechanism involves reversible conversion between two chemical states, the atomically dispersed Pt²⁺ species and sub-nanometer-sized Pt particles in an electrochemical environment. This concept rationalizes the high stability of the Pt–CeO₂ catalyst films with low Pt loading under conditions of dynamically changing electrode potentials.

Regarding the transfer of the above described concepts to other less expensive metals, similar anchoring phenomena were predicted by density functional modeling for various metals from groups 8 to 11. The general trends in the energetics of the corresponding atomically dispersed metals have been established and could for some metals (Pd, Ni) be confirmed by experiments. However, characteristic differences are observed regarding the stability and segregation behavior, rendering the Pt–CeO₂ system unique in this comparison.

Outlook

At present, one challenge regarding the application of single-atom-based Pt–CeO₂ films in fuel cell catalysis is the limited density of the {100} sites, which also limits the total Pt loading and the resulting power density. Different strategies may be employed to tackle this challenge. The first strategy focuses on the improvement of the carbon supports. Along this line, GDLs coated with soot or nitrogen doped carbon films demonstrated great potential.^{45,54} The second strategy involves shaping the morphology of the Pt–CeO₂ film into high-surface area nanostructures with a large fraction of {100} facets, *e.g.* cubes or nanoscrews.^{159,160} This can be achieved, for instance, by innovative deposition techniques or a strain driven growth induced, for example, by doping. As a last strategy, we may consider to prepare the key structure, *i.e.* the square planar PtO₄ moiety, on a different support with a higher degree of nanostructuring.

Additional insight from theoretical modelling studies should also contribute to unravelling in greater detail the mechanism of reversible conversion between Pt²⁺ and the sub-nanometer Pt particles and their interaction with reactants. This will allow establishing a kinetic model of the catalyst under fuel-cell operating conditions, providing valuable understanding of the factors influencing the performance of catalytic materials with atomically dispersed metals.

List of abbreviations

AR XPS	Angle-resolved X-ray photoelectron spectroscopy
nanoGDL, n-GDL	Carbon nanoparticle coated gas diffusion layer

CVD-CNTs	Chemical vapor deposited carbon nanotubes
DFT	Density functional theory
DWCNTs	Double-wall carbon nanotubes
EMSI	Electronic metal–support interactions
GDL	Gas diffusion layer
GLAD	Glancing angle deposition
HAXPS	Hard X-ray photoelectron spectroscopy
SRPES	High-resolution synchrotron spectroscopy
HRTEM	High-resolution transmission electron microscopy
IRAS	Infrared absorption spectroscopy
MWCNTs	Multi-wall carbon nanotubes
ND	Normal deposition
PD	Power density
PEMFC	Proton exchange membrane fuel cell
RER	Resonant enhancement ratio
RPES	Resonant photoemission spectroscopy
SEM	Scanning electron microscopy
SAC	Single atom catalyst
SP	Specific power
UHV	Ultrahigh vacuum
XPS	X-ray photoelectron spectroscopy

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References

- 1 Y. Wang, K. S. Chen, J. Mishler, S. C. Cho and X. C. Adroher, A review of polymer electrolyte membrane fuel cells: Technology, applications, and needs on fundamental research, *Appl. Energy*, 2011, **88**, 981–1007.
- 2 S. C. Kelley, G. A. Deluga and W. H. Smyrl, Miniature fuel cells fabricated on silicon substrates, *AIChE J.*, 2002, **48**, 1071–1082.
- 3 J. D. Morse, Micro-fuel cell power sources, *Int. J. Energy Res.*, 2007, **31**, 576–602.



- 4 T. U. D. o. E. (DOE), Energy Efficiency and Renewable Energy, http://www.eere.energy.gov/hydrogenandfuelcells/mypp/pdfs/fuel_cells.pdf.
- 5 C. Wang, M. Yang and M. Flytzani-Stephanopoulos, Single gold atoms stabilized on nanoscale metal oxide supports are catalytic active centers for various reactions, *AIChE J.*, 2016, **62**, 429–439.
- 6 J. Liu, Catalysis by Supported Single Metal Atoms, *ACS Catal.*, 2017, **7**, 34–59.
- 7 S. Liang, C. Hao and Y. Shi, The Power of Single-Atom Catalysis, *ChemCatChem*, 2015, **7**, 2559–2567.
- 8 X.-F. Yang, A. Wang, B. Qiao, J. Li, J. Liu and T. Zhang, Single-Atom Catalysts: A New Frontier in Heterogeneous Catalysis, *Acc. Chem. Res.*, 2013, **46**, 1740–1748.
- 9 M. Flytzani-Stephanopoulos and B. C. Gates, Atomically Dispersed Supported Metal Catalysts, *Annu. Rev. Chem. Biomol. Eng.*, 2012, **3**, 545–574.
- 10 C. K. Narula and M. Moses-DeBusk, in *Catalysis by Materials with Well-Defined Structures*, ed. S. H. Overbury, Elsevier, Amsterdam, 2015, pp. 263–274.
- 11 N. Cheng, S. Stambula, D. Wang, M. N. Banis, J. Liu, A. Riese, B. Xiao, R. Li, T.-K. Sham, L.-M. Liu, G. A. Botton and X. Sun, Platinum single-atom and cluster catalysis of the hydrogen evolution reaction, *Nat. Commun.*, 2016, **7**, 13638.
- 12 S. Sun, G. Zhang, N. Gauquelin, N. Chen, J. Zhou, S. Yang, W. Chen, X. Meng, D. Geng, M. N. Banis, R. Li, S. Ye, S. Knights, G. A. Botton, T.-K. Sham and X. Sun, Single-atom Catalysis Using Pt/Graphene Achieved through Atomic Layer Deposition, *Sci. Rep.*, 2013, **3**, 1775.
- 13 H. Wang, Q. Feng, Y. Cheng, Y. Yao, Q. Wang, K. Li, U. Schwingenschlöggl, X. X. Zhang and W. Yang, Atomic Bonding between Metal and Graphene, *J. Phys. Chem. C*, 2013, **117**, 4632–4638.
- 14 S. Yang, J. Kim, Y. J. Tak, A. Soon and H. Lee, Single-Atom Catalyst of Platinum Supported on Titanium Nitride for Selective Electrochemical Reactions, *Angew. Chem., Int. Ed.*, 2016, **55**, 2058–2062.
- 15 R.-Q. Zhang, T.-H. Lee, B.-D. Yu, C. Stampfl and A. Soon, The role of titanium nitride supports for single-atom platinum-based catalysts in fuel cell technology, *Phys. Chem. Chem. Phys.*, 2012, **14**, 16552–16557.
- 16 K. Mao, L. Li, W. Zhang, Y. Pei, X. C. Zeng, X. Wu and J. Yang, A Theoretical Study of Single-Atom Catalysis of CO Oxidation Using Au Embedded 2D h-BN Monolayer: A CO-Promoted O₂ Activation, *Sci. Rep.*, 2014, **4**, 5441.
- 17 G. Vilé, D. Albani, M. Nachttegaal, Z. Chen, D. Dontsova, M. Antonietti, N. López and J. Pérez-Ramírez, A Stable Single-Site Palladium Catalyst for Hydrogenations, *Angew. Chem., Int. Ed.*, 2015, **54**, 11265–11269.
- 18 J. M. Thomas, Catalysis: Tens of thousands of atoms replaced by one, *Nature*, 2015, **525**, 325–326.
- 19 J. Lu, C. Aydin, N. D. Browning and B. C. Gates, Imaging Isolated Gold Atom Catalytic Sites in Zeolite NaY, *Angew. Chem., Int. Ed.*, 2012, **51**, 5842–5846.
- 20 A. Uzun, V. Ortalan, Y. Hao, N. D. Browning and B. C. Gates, Imaging Gold Atoms in Site-Isolated MgO-Supported Mononuclear Gold Complexes, *J. Phys. Chem. C*, 2009, **113**, 16847–16849.
- 21 A. Bruix, Y. Lykhach, I. Matolínová, A. Neitzel, T. Skála, N. Tsud, M. Vorokhta, V. Stetsovych, K. Ševčíková, J. Mysliveček, R. Fiala, M. Václavů, K. C. Prince, S. Bruyere, V. Potin, F. Illas, V. Matolín, J. Libuda and K. M. Neyman, Maximum Noble Metal Efficiency in Catalytic Materials: Atomically Dispersed Surface Platinum, *Angew. Chem., Int. Ed.*, 2014, **53**, 10525–10530.
- 22 F. Dvořák, M. Farnesi Camellone, A. Tovt, N.-D. Tran, F. R. Negreiros, M. Vorokhta, T. Skála, I. Matolínová, J. Mysliveček, V. Matolín and S. Fabris, Creating single-atom Pt-ceria catalysts by surface step decoration, *Nat. Commun.*, 2016, **7**, 10801.
- 23 B. T. Qiao, A. Q. Wang, X. F. Yang, L. F. Allard, Z. Jiang, Y. T. Cui, J. Y. Liu, J. Li and T. Zhang, Single-atom catalysis of CO oxidation using Pt₁/FeO_x, *Nat. Chem.*, 2011, **3**, 634–641.
- 24 R. Si and M. Flytzani-Stephanopoulos, Shape and Crystal-Plane Effects of Nanoscale Ceria on the Activity of Au-CeO₂ Catalysts for the Water–Gas Shift Reaction, *Angew. Chem., Int. Ed.*, 2008, **47**, 2884–2887.
- 25 J. H. Kwak, J. Z. Hu, D. Mei, C. W. Yi, D. H. Kim, C. H. F. Peden, L. F. Allard and J. Szanyi, Coordinatively unsaturated Al³⁺ centers as binding sites for active catalyst phases of platinum on gamma-Al₂O₃, *Science*, 2009, **325**, 1670–1673.
- 26 P. Hu, Z. Huang, Z. Amghouz, M. Makkee, F. Xu, F. Kapteijn, A. Dikhtiarenko, Y. Chen, X. Gu and X. Tang, Electronic Metal–Support Interactions in Single-Atom Catalysts, *Angew. Chem., Int. Ed.*, 2014, **53**, 3418–3421.
- 27 J. Jones, H. Xiong, A. T. DeLaRiva, E. J. Peterson, H. Pham, S. R. Challa, G. Qi, S. Oh, M. H. Wiebenga, X. I. Pereira Hernández, Y. Wang and A. K. Datye, Thermally stable single-atom platinum-on-ceria catalysts via atom trapping, *Science*, 2016, **353**, 150–154.
- 28 L. Wang, W. Zhang, S. Wang, Z. Gao, Z. Luo, X. Wang, R. Zeng, A. Li, H. Li, M. Wang, X. Zheng, J. Zhu, W. Zhang, C. Ma, R. Si and J. Zeng, Atomic-level insights in optimizing reaction paths for hydroformylation reaction over Rh/CoO single-atom catalyst, *Nat. Commun.*, 2016, **7**, 14036.
- 29 T. Rajesh, A. K. Rajarajan, C. S. Gopinath and R. N. Devi, Evidence of Cationic Pt Active for Water-Gas Shift Reaction: Pt-Doped BaCeO₃ Perovskite, *J. Phys. Chem. C*, 2012, **116**, 9526–9532.
- 30 J. Liu, Advanced Electron Microscopy of Metal–Support Interactions in Supported Metal Catalysts, *ChemCatChem*, 2011, **3**, 934–948.
- 31 G. S. Parkinson, Z. Novotny, G. Argentero, M. Schmid, J. Pavelec, R. Kosak, P. Blaha and U. Diebold, Carbon monoxide-induced adatom sintering in a Pd–Fe₃O₄ model catalyst, *Nat. Mater.*, 2013, **12**, 724–728.
- 32 K. T. Rim, D. Eom, L. Liu, E. Stolyarova, J. M. Raitano, S.-W. Chan, M. Flytzani-Stephanopoulos and G. W. Flynn, Charging and Chemical Reactivity of Gold Nanoparticles and Adatoms on the (111) Surface of Single-Crystal Magnetite: A Scanning Tunneling Microscopy/Spectroscopy Study, *J. Phys. Chem. C*, 2009, **113**, 10198–10205.



- 33 Z. Huang, X. Gu, Q. Cao, P. Hu, J. Hao, J. Li and X. Tang, Catalytically Active Single-Atom Sites Fabricated from Silver Particles, *Angew. Chem., Int. Ed.*, 2012, 51, 4198–4203.
- 34 C. T. Campbell, Catalyst-support interactions: Electronic perturbations, *Nat. Chem.*, 2012, 4, 597–598.
- 35 G. Pacchioni, Electronic interactions and charge transfers of metal atoms and clusters on oxide surfaces, *Phys. Chem. Chem. Phys.*, 2013, 15, 1737–1757.
- 36 M. S. Hegde and P. Bera, Noble Metal Ion Substituted CeO₂ Catalysts: Electronic Interaction between Noble Metal Ions and CeO₂ Lattice, *Catal. Today*, 2015, 253, 40–50.
- 37 M. S. Hegde, G. Madras and K. C. Patil, Noble Metal Ionic Catalysts, *Acc. Chem. Res.*, 2009, 42, 704–712.
- 38 E. W. McFarland and H. Metiu, Catalysis by Doped Oxides, *Chem. Rev.*, 2013, 113, 4391–4427.
- 39 V. Matolín, M. Cabala, I. Matolínová, M. Škoda, M. Václavů, K. C. Prince, T. Skála, T. Mori, H. Yoshikawa, Y. Yamashita, S. Ueda and K. Kobayashi, Pt and Sn Doped Sputtered CeO₂ Electrodes for Fuel Cell Applications, *Fuel Cells*, 2010, 10, 139–144.
- 40 M. Vaclavu, I. Matolinova, J. Myslivecek, R. Fiala and V. Matolin, Anode Material for Hydrogen Polymer Membrane Fuel Cell: Pt-CeO₂ RF-Sputtered Thin Films, *J. Electrochem. Soc.*, 2009, 156, B938–B942.
- 41 R. Fiala, A. Figueroba, A. Bruix, M. Vaclavu, A. Rednyk, I. Khalakhan, M. Vorokhta, J. Lavkova, F. Illas, V. Potin, I. Matolinova, K. M. Neyman and V. Matolin, High efficiency of Pt²⁺-CeO₂ novel thin film catalyst as anode for proton exchange membrane fuel cells, *Appl. Catal., B*, 2016, 197, 262–270.
- 42 V. Matolín, I. Matolínová, M. Václavů, I. Khalakhan, M. Vorokhta, R. Fiala, I. Piš, Z. Sofer, J. Poltiová-Vejpravová, T. Mori, V. Potin, H. Yoshikawa, S. Ueda and K. Kobayashi, Platinum-doped CeO₂ thin film catalysts prepared by magnetron sputtering, *Langmuir*, 2010, 26, 12824–12831.
- 43 R. Fiala, I. Khalakhan, I. Matolínová, M. Václavů, M. Vorokhta, Z. Sofer, S. Huber, V. Potin and V. Matolín, Pt-CeO₂ Coating of Carbon Nanotubes Grown on Anode Gas Diffusion Layer of the Polymer Electrolyte Membrane Fuel Cell, *J. Nanosci. Nanotechnol.*, 2011, 11, 5062–5067.
- 44 R. Fiala, M. Vaclavu, A. Rednyk, I. Khalakhan, M. Vorokhta, J. Lavkova, V. Potin, I. Matolinova and V. Matolin, Pt–CeO_x Thin Film Catalysts for PEMFC, *Catal. Today*, 2015, 240(Part B), 236–241.
- 45 I. Khalakhan, R. Fiala, J. Lavková, P. Kúš, A. Ostroverkh, M. Václavů, M. Vorokhta, I. Matolínová and V. Matolín, Candle Soot as Efficient Support for Proton Exchange Membrane Fuel Cell Catalyst, *Fuel Cells*, 2016, 16, 652–655.
- 46 V. Matolín, I. Khalakhan, I. Matolínová, M. Václavů, K. Veltruská and M. Vorokhta, Pt^{2+,4+} ions in CeO₂ rf-sputtered thin films, *Surf. Interface Anal.*, 2010, 42, 882–885.
- 47 R. Fiala, M. Vaclavu, M. Vorokhta, I. Khalakhan, J. Lavková, V. Potin, I. Matolínová and V. Matolín, Proton Exchange Membrane Fuel Cell Made of Magnetron Sputtered Pt-CeO_x and Pt-Co Thin Film Catalysts, *J. Power Sources*, 2015, 273, 105–109.
- 48 H. Yoshikawa, I. Matolínová and V. Matolín, Practical chemical analysis of Pt and Pd based heterogeneous catalysts with hard X-ray photoelectron spectroscopy, *J. Electron Spectrosc. Relat. Phenom.*, 2013, 190(Part B), 268–277.
- 49 K. Ševčíková, V. Nehasil, M. Vorokhta, S. Haviar, V. Matolín, I. Matolínová, K. Mašek, I. Piš, K. Kobayashi, M. Kobata, T. Nagata, Y. Matsushita and H. Yoshikawa, Altering properties of cerium oxide thin films by Rh doping, *Mater. Res. Bull.*, 2015, 67, 5–13.
- 50 V. Matolin, M. Cabala, I. Matolinova, M. Skoda, J. Libra, M. Vaclavu, K. C. Prince, T. Skala, H. Yoshikawa, Y. Yamashita, S. Ueda and K. Kobayashi, Au⁺ and Au³⁺ ions in CeO₂ rf-sputtered thin films, *J. Phys. D: Appl. Phys.*, 2009, 42, 115301.
- 51 V. Matolín, R. Fiala, I. Khalakhan, J. Lavková, M. Václavů and M. Vorokhta, Nanoporous Ptⁿ⁺-CeO_x catalyst films grown on carbon substrates, *Int. J. Nanotechnol.*, 2012, 9, 680–694.
- 52 M. Vorokhta, I. Khalakhan, I. Matolínová, M. Kobata, H. Yoshikawa, K. Kobayashi and V. Matolín, Nanostructured Pt–CeO₂ thin film catalyst grown on graphite foil by magnetron sputtering, *Appl. Surf. Sci.*, 2013, 267, 119–123.
- 53 I. Khalakhan, M. Dubau, S. Haviar, J. Lavková, I. Matolínová, V. Potin, M. Vorokhta and V. Matolín, Growth of nano-porous Pt-doped cerium oxide thin films on glassy carbon substrate, *Ceram. Int.*, 2013, 39, 3765–3769.
- 54 M. Dubau, J. Lavková, I. Khalakhan, S. Haviar, V. Potin, V. R. Matolín and I. Matolínová, Preparation of Magnetron Sputtered Thin Cerium Oxide Films with a Large Surface on Silicon Substrates Using Carbonaceous Interlayers, *ACS Appl. Mater. Interfaces*, 2014, 6, 1213–1218.
- 55 S. Haviar, M. Dubau, J. Lavková, I. Khalakhan, V. Potin, V. Matolín and I. Matolínová, Investigation of Growth Mechanism of Thin Sputtered Cerium Oxide Films on Carbon Substrates, *Sci. Adv. Mater.*, 2014, 6, 1278–1285.
- 56 J. Lavkova, I. Khalakhan, M. Chundak, M. Vorokhta, V. Potin, V. Matolin and I. Matolinova, Growth and composition of nanostructured and nanoporous cerium oxide thin films on a graphite foil, *Nanoscale*, 2015, 7, 4038–4047.
- 57 M. Vorokhta, I. Khalakhan, I. Matolínová, J. Nováková, S. Haviar, J. Lančok, M. Novotný, H. Yoshikawa and V. Matolín, PLD prepared nanostructured Pt-CeO₂ thin films containing ionic platinum, *Appl. Surf. Sci.*, 2017, 396, 278–283.
- 58 N. Zanfoni, L. Avril, L. Imhoff, B. Domenichini and S. Bourgeois, Direct liquid injection chemical vapor deposition of platinum doped cerium oxide thin films, *Thin Solid Films*, 2015, 589, 246–251.
- 59 L. Avril, N. Zanfoni, P. Simon, L. Imhoff, S. Bourgeois and B. Domenichini, MOCVD growth of porous cerium oxide thin films on silicon substrate, *Surf. Coat. Technol.*, 2015, 280, 148–153.



- 60 P. Simon, N. Zanfoni, L. Avril, Z. Li, V. Potin, B. Domenichini and S. Bourgeois, Nanoporous Platinum Doped Cerium Oxides Thin Films Grown on Silicon Substrates: Ionic Platinum Localization and Stability, *Adv. Mater. Interfaces*, 2017, 4, 1600821.
- 61 P. Simon, N. Zanfoni, L. Avril, J. Lavkova, I. Matolinova, L. Imhoff, V. Potin, B. Domenichini and S. Bourgeois, Observation of surface reduction in porous ceria thin film grown on graphite foil substrate, *Mater. Today*, 2016, 3, 2772–2779.
- 62 I. Matolinová, R. Fiala, I. Khalakhan, M. Vorokhta, Z. Sofer, H. Yoshikawa, K. Kobayashi and V. Matolín, Synchrotron radiation photoelectron spectroscopy study of metal-oxide thin film catalysts: Pt-CeO₂ coated CNTs, *Appl. Surf. Sci.*, 2012, 258, 2161–2164.
- 63 S. Bruyère, A. Caccucci, V. Potin, I. Matolinova, M. Vorokhta and V. Matolin, Deposition of Pt and Sn doped CeOx layers on silicon substrate, *Surf. Coat. Technol.*, 2013, 227, 15–18.
- 64 V. Potin, J. Lavkova, S. Bourgeois, M. Dubau, I. Matolinova and V. Matolin, Structural and Chemical Characterization of Cerium Oxide Thin Layers Grown on Silicon Substrate, *Mater. Today*, 2015, 2, 101–107.
- 65 A. Migani, G. N. Vayssilov, S. T. Bromley, F. Illas and K. M. Neyman, Dramatic Reduction of the Oxygen Vacancy Formation Energy in Ceria Particles: A Possible Key to Their Remarkable Reactivity at the Nanoscale, *J. Mater. Chem.*, 2010, 20, 10535–10546.
- 66 A. Migani, K. M. Neyman and S. T. Bromley, Octahedrality versus tetrahedrality in stoichiometric ceria nanoparticles, *Chem. Commun.*, 2012, 48, 4199–4201.
- 67 F. Zhang, Q. Jin and S.-W. Chan, Ceria nanoparticles: Size, size distribution, and shape, *J. Appl. Phys.*, 2004, 95, 4319–4326.
- 68 A. Migani, G. N. Vayssilov, S. T. Bromley, F. Illas and K. M. Neyman, Greatly Facilitated Oxygen Vacancy Formation in Ceria Nanocrystallites, *Chem. Commun.*, 2010, 46, 5936–5938.
- 69 M. M. Branda, R. M. Ferullo, M. Causà and F. Illas, Relative stabilities of low index and stepped CeO₂ surfaces from hybrid and GGA plus U implementations of density functional theory, *J. Phys. Chem. C*, 2011, 115, 3716–3721.
- 70 A. Bruix and K. M. Neyman, Modeling Ceria-Based Nanomaterials for Catalysis and Related Applications, *Catal. Lett.*, 2016, 146, 2053–2080.
- 71 A. Bruix, K. M. Neyman and F. Illas, Adsorption, oxidation state, and diffusion of Pt atoms on the CeO₂(111) surface, *J. Phys. Chem. C*, 2010, 114, 14202–14207.
- 72 A. Bruix, F. Nazari, K. M. Neyman and F. Illas, On the adsorption and formation of Pt dimers on the CeO₂(111) surface, *J. Chem. Phys.*, 2011, 135, 244708.
- 73 A. Bruix, A. Migani, G. N. Vayssilov, K. M. Neyman, J. Libuda and F. Illas, Effects of deposited Pt particles on the reducibility of CeO₂(111), *Phys. Chem. Chem. Phys.*, 2011, 13, 11384–11392.
- 74 K. M. Neyman, C. Inntam, V. A. Nasluzov, R. Kosarev and N. Rösch, Adsorption of d-metal atoms on the regular MgO(001) surface: Density functional study of cluster models embedded in an elastic polarizable environment, *Appl. Phys. A: Mater. Sci. Process.*, 2004, 78, 823–828.
- 75 K. M. Neyman, C. Inntam, A. V. Matveev, V. A. Nasluzov and N. Rösch, Single d-Metal Atoms on F_s and F_s⁺ Defects of MgO(001): A Theoretical Study across the Periodic Table, *J. Am. Chem. Soc.*, 2005, 127, 11652–11660.
- 76 A. G. Hu, K. M. Neyman, M. Staufer, T. Belling, B. C. Gates and N. Rosch, A Surface Site as Polydentate Ligand of a Metal Complex: Density Functional Studies of Rhenium Subcarbonyls Supported on Magnesium Oxide, *J. Am. Chem. Soc.*, 1999, 121, 4522–4523.
- 77 F. A. Cotton and G. Wilkinson, *Advanced Inorganic Chemistry*, Wiley, New York, 1990.
- 78 C. Kittel, *Introduction to Solid State Physics*, Wiley, New York, 1976.
- 79 Y. Nagai, T. Hirabayashi, K. Dohmae, N. Takagi, T. Minami, H. Shinjoh and S. Matsumoto, Sintering inhibition mechanism of platinum supported on ceria-based oxide and Pt-oxide-support interaction, *J. Catal.*, 2006, 242, 103–109.
- 80 N. Nilius, S. M. Kozlov, J.-F. Jerratsch, M. Baron, X. Shao, F. Viñes, S. Shaikhutdinov, K. M. Neyman and H.-J. Freund, Formation of One-Dimensional Electronic States along the Step Edges of CeO₂(111), *ACS Nano*, 2012, 6, 1126–1133.
- 81 S. M. Kozlov, F. Viñes, N. Nilius, S. Shaikhutdinov and K. M. Neyman, Absolute Surface Step Energies: Accurate Theoretical Methods Applied to Ceria Nanoislands, *J. Phys. Chem. Lett.*, 2012, 3, 1956–1961.
- 82 A. Figueroba, G. Kovács, A. Bruix and K. M. Neyman, Towards stable single-atom catalysts: Strong binding of atomically dispersed transition metals on the surface of nanostructured ceria, *Catal. Sci. Technol.*, 2016, 6, 6806–6813.
- 83 M. A. Sk, S. M. Kozlov, K. H. Lim, A. Migani and K. M. Neyman, Oxygen Vacancies in Self-Assemblies of Ceria Nanoparticles, *J. Mater. Chem. A*, 2014, 2, 18329–18338.
- 84 R. Nedyalkova, D. Niznansky and A.-C. Roger, Iron–ceria–zirconia fluorite catalysts for methane selective oxidation to formaldehyde, *Catal. Commun.*, 2009, 10, 1875–1880.
- 85 A. Satsuma, M. Yanagihara, J. Ohyama and K. Shimizu, Oxidation of CO over Ru/Ceria prepared by self-dispersion of Ru metal powder into nano-sized particle, *Catal. Today*, 2013, 201, 62–67.
- 86 S. S. Y. Lin, D. H. Kim, M. H. Engelhard and S. Y. Ha, Water-induced formation of cobalt oxides over supported cobalt/ceria–zirconia catalysts under ethanol-steam conditions, *J. Catal.*, 2010, 273, 229–235.
- 87 M. Haneda, K. Shinoda, A. Nagane, O. Houshito, H. Takagi, Y. Nakahara, K. Hiroe, T. Fujitani and H. Hamada, Catalytic performance of rhodium supported on ceria–zirconia mixed oxides for reduction of NO by propene, *J. Catal.*, 2008, 259, 223–231.



- 88 Y. Huang, A. Wang, L. Li, X. Wang, D. Su and T. Zhang, "Ir-in-ceria": A highly selective catalyst for preferential CO oxidation, *J. Catal.*, 2008, **255**, 144–152.
- 89 H. A. Aleksandrov, K. M. Neyman and G. N. Vayssilov, The Structure and Stability of Reduced and Oxidized Mononuclear Platinum Species on Nanostructured Ceria from Density Functional Modeling, *Phys. Chem. Chem. Phys.*, 2015, **17**, 14551–14560.
- 90 T. Wu, X. Pan, Y. Zhang, Z. Miao, B. Zhang, J. Li and X. Yang, Investigation of the Redispersion of Pt Nanoparticles on Polyhedral Ceria Nanoparticles, *J. Phys. Chem. Lett.*, 2014, **5**, 2479–2483.
- 91 Y. Gao, W. Wang, S. Chang and W. Huang, Morphology Effect of CeO₂ Support in the Preparation, Metal–Support Interaction, and Catalytic Performance of Pt/CeO₂ Catalysts, *ChemCatChem*, 2013, **5**, 3610–3620.
- 92 J. Carrasco, D. López-Durán, Z. Liu, T. Duchoň, J. Evans, S. D. Senanayake, E. J. Crumlin, V. Matolín, J. A. Rodríguez and M. V. Ganduglia-Pirovano, In Situ and Theoretical Studies for the Dissociation of Water on an Active Ni/CeO₂ Catalyst: Importance of Strong Metal–Support Interactions for the Cleavage of O–H Bonds, *Angew. Chem., Int. Ed.*, 2015, **54**, 3917–3921.
- 93 S. Colussi, A. Gayen, M. Farnesi Camellone, M. Boaro, J. Llorca, S. Fabris and A. Trovarelli, Nanofaceted Pd–O Sites in Pd–Ce Surface Superstructures: Enhanced Activity in Catalytic Combustion of Methane, *Angew. Chem., Int. Ed.*, 2009, **48**, 8481–8484.
- 94 R. V. Gulyaev, T. Y. Kardash, S. E. Malykhin, O. A. Stonkus, A. S. Ivanova and A. I. Boronin, The local structure of Pd_xCe_{1-x}O_{2-xδ} solid solutions, *Phys. Chem. Chem. Phys.*, 2014, **16**, 13523–13539.
- 95 J. S. Elias, M. Risch, L. Giordano, A. N. Mansour and Y. Shao-Horn, Structure, Bonding, and Catalytic Activity of Monodisperse, Transition-Metal-Substituted CeO₂ Nanoparticles, *J. Am. Chem. Soc.*, 2014, **136**, 17193–17200.
- 96 H. Müller-Buschbaum, Zur Kristallchemie der Oxoargentate und Silberoxometallate, *Z. Anorg. Allg. Chem.*, 2004, **630**, 2125–2175.
- 97 A. M. Venezia, G. Pantaleo, A. Longo, G. Di Carlo, M. P. Casaletto, F. L. Liotta and G. Deganello, Relationship between Structure and CO Oxidation Activity of Ceria-Supported Gold Catalysts, *J. Phys. Chem. B*, 2005, **109**, 2821–2827.
- 98 A. Neitzel, A. Figueroba, Y. Lykhach, T. Skála, M. Vorokhta, N. Tsud, S. Mehl, K. Ševčíková, K. C. Prince, K. M. Neyman, V. Matolín and J. Libuda, Atomically Dispersed Pd, Ni and Pt Species in Ceria-Based Catalysts: Principal Differences in Stability and Reactivity, *J. Phys. Chem. C*, 2016, **120**, 9852–9862.
- 99 Y. Lykhach, A. Figueroba, M. F. Camellone, A. Neitzel, T. Skála, F. R. Negreiros, M. Vorokhta, N. Tsud, K. C. Prince, S. Fabris, K. M. Neyman, V. Matolín and J. Libuda, Reactivity of Atomically Dispersed Pt²⁺ Species towards H₂: Model Pt–CeO₂ Fuel Cell Catalyst, *Phys. Chem. Chem. Phys.*, 2016, **18**, 7672–7679.
- 100 F. Dvořák, O. Stetsovych, M. Steger, E. Cherradi, I. Matolínová, N. Tsud, M. Škoda, T. Skála, J. Mysliveček and V. Matolín, Adjusting Morphology and Surface Reduction of CeO₂(111) Thin Films on Cu(111), *J. Phys. Chem. C*, 2011, **115**, 7496–7503.
- 101 D. R. Mullins, P. M. Albrecht, T.-L. Chen, F. C. Calaza, M. D. Biegalski, H. M. Christen and S. H. Overbury, Water Dissociation on CeO₂(100) and CeO₂(111) Thin Films, *J. Phys. Chem. C*, 2012, **116**, 19419–19428.
- 102 A. Bruix, J. A. Rodríguez, P. J. Ramírez, S. D. Senanayake, J. Evans, J. B. Park, D. Stacchiola, P. Liu, J. Hrbek and F. Illas, A New Type of Strong Metal–Support Interaction and the Production of H₂ through the Transformation of Water on Pt/CeO₂(111) and Pt/CeO_x/TiO₂(110) Catalysts, *J. Am. Chem. Soc.*, 2012, **134**, 8968–8974.
- 103 Q. Fu, H. Saltsburg and M. Flytzani-Stephanopoulos, Active Nonmetallic Au and Pt Species on Ceria-Based Water-Gas Shift Catalysts, *Science*, 2003, **301**, 935–938.
- 104 M. Hatanaka, N. Takahashi, T. Tanabe, Y. Nagai, K. Dohmae, Y. Aoki, T. Yoshida and H. Shinjoh, Ideal Pt loading for a Pt/CeO₂-based catalyst stabilized by a Pt–O–Ce bond, *Appl. Catal., B*, 2010, **99**, 336–342.
- 105 G. Ertl, H. Knözinger, F. Schüth and J. Weitkamp, *Handbook of Heterogeneous Catalysis*, Wiley, Weinheim, 2008.
- 106 F. M. Hoffmann, Infrared reflection-absorption spectroscopy of adsorbed molecules, *Surf. Sci. Rep.*, 1983, **3**, 107–192.
- 107 J. M. Chalmers and P. R. Griffiths, *Handbook of vibrational spectroscopy*, John Wiley & Sons, Ltd., Chichester, UK, 2001.
- 108 K. I. Hadjiivanov and G. N. Vayssilov, in *Adv. Catal.*, Academic Press, 2002, vol. 47, pp. 307–511.
- 109 O. Brummel, F. Waidhas, F. Faisal, R. Fiala, M. Vorokhta, I. Khalakhan, M. Dubau, A. Figueroba, G. Kovács, H. A. Aleksandrov, G. N. Vayssilov, S. M. Kozlov, K. M. Neyman, V. Matolín and J. Libuda, Stabilization of Small Platinum Nanoparticles on Pt–CeO₂ Thin Film Electrocatalysts During Methanol Oxidation, *J. Phys. Chem. C*, 2016, **120**, 19723–19736.
- 110 H. A. Aleksandrov, K. M. Neyman, K. I. Hadjiivanov and G. N. Vayssilov, Can the state of platinum species be unambiguously determined by the stretching frequency of an adsorbed CO probe molecule?, *Phys. Chem. Chem. Phys.*, 2016, **18**, 22108–22121.
- 111 M. J. Kappers and J. H. van der Maas, Correlation between CO frequency and Pt coordination number. A DRIFT study on supported Pt catalysts, *Catal. Lett.*, 1991, **10**, 365–373.
- 112 A. Neitzel, Y. Lykhach, T. Skála, N. Tsud, M. Vorokhta, D. Mazur, K. C. Prince, V. Matolín and J. Libuda, Surface Sites on Pt–CeO₂ Mixed Oxide Catalysts Probed by CO Adsorption: A Synchrotron Radiation Photoelectron Spectroscopy Study, *Phys. Chem. Chem. Phys.*, 2014, **16**, 24747–24754.
- 113 G. N. Vayssilov, M. Mihaylov, P. S. Petkov, K. I. Hadjiivanov and K. M. Neyman, Reassignment of the Vibrational



- Spectra of Carbonates, Formates, and Related Surface Species on Ceria: A Combined Density Functional and Infrared Spectroscopy Investigation, *J. Phys. Chem. C*, 2011, **115**, 23435–23454.
- 114 G. Dutta, U. V. Waghmare, T. Baidya and M. S. Hegde, Hydrogen Spillover on CeO₂/Pt: Enhanced Storage of Active Hydrogen, *Chem. Mater.*, 2007, **19**, 6430–6436.
- 115 P. Bera, K. R. Priolkar, A. Gayen, P. R. Sarode, M. S. Hegde, S. Emura, R. Kumashiro, V. Jayaram and G. N. Subbanna, Ionic Dispersion of Pt over CeO₂ by the Combustion Method: Structural Investigation by XRD, TEM, XPS, and EXAFS, *Chem. Mater.*, 2003, **15**, 2049–2060.
- 116 B. D. Mukri, U. V. Waghmare and M. S. Hegde, Platinum Ion-Doped TiO₂: High Catalytic Activity of Pt²⁺ with Oxide Ion Vacancy in Ti⁴⁺_{1-x}Pt²⁺_xO_{2-x} Compared to Pt⁴⁺ without Oxide Ion Vacancy in Ti⁴⁺_{1-x}Pt⁴⁺_xO₂, *Chem. Mater.*, 2013, **25**, 3822–3833.
- 117 Y. Lykhach, S. M. Kozlov, T. Skála, A. Tovt, V. Stetsovych, N. Tsud, F. Dvořák, V. Johánek, A. Neitzel, J. Mysliveček, S. Fabris, V. Matolín, K. M. Neyman and J. Libuda, Counting Electrons on Supported Nanoparticles, *Nat. Mater.*, 2016, **15**, 284–288.
- 118 Y. Lykhach, T. Staudt, M. Vorokhta, T. Skála, V. Johánek, K. C. Prince, V. Matolín and J. Libuda, Hydrogen Spillover Monitored by Resonant Photoemission Spectroscopy, *J. Catal.*, 2012, **285**, 6–9.
- 119 G. N. Vayssilov, Y. Lykhach, A. Migani, T. Staudt, G. P. Petrova, N. Tsud, T. Skála, A. Bruix, F. Illas, K. C. Prince, V. Matolín, K. M. Neyman and J. Libuda, Support Nanostructure Boosts Oxygen Transfer to Catalytically Active Platinum Nanoparticles, *Nat. Mater.*, 2011, **10**, 310–315.
- 120 C. Kittel, *Introduction to Solid State Physics, 8th edition*, Wiley, 2004.
- 121 A. Neitzel, V. Johánek, Y. Lykhach, T. Skála, N. Tsud, M. Vorokhta, V. Matolín and J. Libuda, Reduction of Pt²⁺ species in model Pt–CeO₂ fuel cell catalysts upon reaction with methanol, *Appl. Surf. Sci.*, 2016, **387**, 674–681.
- 122 V. Matolín, J. Libra, M. Škoda, N. Tsud, K. C. Prince and T. Skála, Methanol Adsorption on a CeO₂(111)/Cu(111) Thin Film Model Catalyst, *Surf. Sci.*, 2009, **603**, 1087–1092.
- 123 D. R. Mullins, M. D. Robbins and J. Zhou, Adsorption and Reaction of Methanol on Thin-Film Cerium Oxide, *Surf. Sci.*, 2006, **600**, 1547–1558.
- 124 M. Capdevila-Cortada, M. García-Melchor and N. López, Unraveling the structure sensitivity in methanol conversion on CeO₂: A DFT + U study, *J. Catal.*, 2015, **327**, 58–64.
- 125 D. J. Miller, H. Öberg, S. Kaya, H. Sanchez Casalongue, D. Friebel, T. Anniyev, H. Ogasawara, H. Bluhm, L. G. M. Pettersson and A. Nilsson, Oxidation of Pt(111) under Near-Ambient Conditions, *Phys. Rev. Lett.*, 2011, **107**, 195502.
- 126 F. Pilger, A. Testino, A. Carino, C. Proff, A. Kambolis, A. Cervellino and C. Ludwig, Size Control of Pt Clusters on CeO₂ Nanoparticles via an Incorporation–Segregation Mechanism and Study of Segregation Kinetics, *ACS Catal.*, 2016, **6**, 3688–3699.
- 127 Y. Lykhach, A. Figueroba, T. Skála, T. Duchoň, N. Tsud, M. Aulická, A. Neitzel, K. Veltruská, K. C. Prince, V. Matolín, K. M. Neyman and J. Libuda, Redox-mediated conversion of atomically dispersed platinum to sub-nanometer particles, *J. Mater. Chem. A*, 2017, **5**, 9250–9261.
- 128 J. Paier, C. Penschke and J. Sauer, Oxygen Defects and Surface Chemistry of Ceria: Quantum Chemical Studies Compared to Experiment, *Chem. Rev.*, 2013, **113**, 3949–3985.
- 129 M. V. Ganduglia-Pirovano, J. L. F. Da Silva and J. Sauer, Density-Functional Calculations of the Structure of Near-Surface Oxygen Vacancies and Electron Localization on CeO₂(111), *Phys. Rev. Lett.*, 2009, **102**, 026101–026104.
- 130 T. Skála, F. Šutara, K. C. Prince and V. Matolín, Cerium oxide stoichiometry alteration via Sn deposition: Influence of temperature, *J. Electron Spectrosc. Relat. Phenom.*, 2009, **169**, 20–25.
- 131 M. Škoda, M. Cabala, V. Cháb, K. C. Prince, L. Sedláček, T. Skála, F. Šutara and V. Matolín, Sn interaction with CeO₂(111) system: Bimetallic bonding and ceria reduction, *Appl. Surf. Sci.*, 2008, **254**, 4375–4379.
- 132 A. Neitzel, Y. Lykhach, T. Skála, N. Tsud, V. Johánek, M. Vorokhta, K. C. Prince, V. Matolín and J. Libuda, Hydrogen activation on Pt–Sn nanoalloys supported on mixed Sn–Ce oxide films, *Phys. Chem. Chem. Phys.*, 2014, **16**, 13209–13219.
- 133 C. T. Campbell, S. C. Parker and D. E. Starr, The Effect of Size-Dependent Nanoparticle Energetics on Catalyst Sintering, *Science*, 2002, **298**, 811–814.
- 134 G. N. Vayssilov, A. Migani and K. Neyman, Density Functional Modeling of the Interactions of Platinum Clusters with CeO₂ Nanoparticles of Different Size, *J. Phys. Chem. C*, 2011, **115**, 16081–16086.
- 135 Z. Lu and Z. Yang, Interfacial properties of NM/CeO₂(111) (NM = noble metal atoms or clusters of Pd, Pt and Rh): a first principles study, *J. Phys.: Condens. Matter*, 2010, **22**, 475003.
- 136 S. M. Kozlov and K. M. Neyman, Effects of electron transfer in model catalysts composed of Pt nanoparticles on CeO₂(111) surface, *J. Catal.*, 2016, **344**, 507–514.
- 137 F. R. Negreiros and S. Fabris, Role of Cluster Morphology in the Dynamics and Reactivity of Subnanometer Pt Clusters Supported on Ceria Surfaces, *J. Phys. Chem. C*, 2014, **118**, 21014–21020.
- 138 Y. Pan, N. Nilius, H.-J. Freund, J. Paier, C. Penschke and J. Sauer, Titration of Ce³⁺ Ions in the CeO₂(111) Surface by Au Adatoms, *Phys. Rev. Lett.*, 2013, **111**, 206101.
- 139 K. Ševčíková, L. Szabová, M. Kettner, P. Homola, N. Tsud, S. Fabris, V. Matolín and V. Nehasil, Experimental and Theoretical Study on the Electronic Interaction between Rh Adatoms and CeO_x Substrate in Dependence on a Degree of Cerium Oxide Reduction, *J. Phys. Chem. C*, 2016, **120**, 5468–5476.
- 140 K. Ševčíková, T. Kolářová, T. Skála, N. Tsud, M. Václavů, Y. Lykhach, V. Matolín and V. Nehasil, Impact of Rh–CeO_x interaction on CO oxidation mechanisms, *Appl. Surf. Sci.*, 2015, **332**, 747–755.



- 141 Y. Lykhach, V. Johánek, H. Aleksandrov, S. M. Kozlov, M. Happel, T. Skála, P. S. Petkov, N. Tsud, G. N. Vayssilov, K. C. Prince, K. M. Neyman, V. Matolín and J. Libuda, Water Chemistry on Model Ceria and Pt/Ceria Catalysts, *J. Phys. Chem. C*, 2012, **116**, 12103–12113.
- 142 V. Matolín, I. Matolínová, F. Dvořák, V. Johánek, J. Mysliveček, K. C. Prince, T. Skála, O. Stetsovyeh, N. Tsud, M. Václavů and B. Šmíd, Water Interaction with CeO₂(111)/Cu(111) Model Catalyst Surface, *Catal. Today*, 2012, **181**, 124–132.
- 143 L. Szabová, Y. Tateyama, V. Matolín and S. Fabris, Water Adsorption and Dissociation at Metal-Supported Ceria Thin Films: Thickness and Interface-Proximity Effects Studied with DFT+U Calculations, *J. Phys. Chem. C*, 2015, **119**, 2537–2544.
- 144 M. Farnesi Camellone, F. Negreiros Ribeiro, L. Szabová, Y. Tateyama and S. Fabris, Catalytic Proton Dynamics at the Water/Solid Interface of Ceria-Supported Pt Clusters, *J. Am. Chem. Soc.*, 2016, **138**, 11560–11567.
- 145 S. M. Kozlov and K. M. Neyman, O vacancies on steps on the CeO₂(111) surface, *Phys. Chem. Chem. Phys.*, 2014, **16**, 7823–7829.
- 146 S. Carrettin, P. Concepción, A. Corma, J. M. López Nieto and V. F. Puntes, Nanocrystalline CeO₂ Increases the Activity of Au for CO Oxidation by Two Orders of Magnitude, *Angew. Chem., Int. Ed.*, 2004, **43**, 2538–2540.
- 147 V. Fiorin, D. Borthwick and D. A. King, Microcalorimetry of O₂ and NO on flat and stepped platinum surfaces, *Surf. Sci.*, 2009, **603**, 1360–1364.
- 148 P. Ghosh, M. Farnesi Camellone and S. Fabris, Fluxionality of Au Clusters at Ceria Surfaces during CO Oxidation: Relationships among Reactivity, Size, Cohesion, and Surface Defects from DFT Simulations, *J. Phys. Chem. Lett.*, 2013, **4**, 2256–2263.
- 149 P. C. Jennings, H. A. Aleksandrov, K. M. Neyman and R. L. Johnston, A DFT study of oxygen dissociation on platinum based nanoparticles, *Nanoscale*, 2014, **6**, 1153–1165.
- 150 P. C. Jennings, H. A. Aleksandrov, K. M. Neyman and R. L. Johnston, DFT studies of oxygen dissociation on the 116-atom platinum truncated octahedron particle, *Phys. Chem. Chem. Phys.*, 2014, **16**, 26539–26545.
- 151 P. C. Jennings, H. A. Aleksandrov, K. M. Neyman and R. L. Johnston, O₂ Dissociation on M@Pt Core–Shell Particles for 3d, 4d, and 5d Transition Metals, *J. Phys. Chem. C*, 2015, **119**, 11031–11041.
- 152 F. Viñes, C. Loschen, F. Illas and K. M. Neyman, Edge sites as a gate for subsurface carbon in palladium nanoparticles, *J. Catal.*, 2009, **266**, 59–63.
- 153 K. M. Neyman and S. Schauermaun, Hydrogen Diffusion into Palladium Nanoparticles: Pivotal Promotion by Carbon, *Angew. Chem., Int. Ed.*, 2010, **49**, 4743–4746.
- 154 M. Hatanaka, N. Takahashi, N. Takahashi, T. Tanabe, Y. Nagai, A. Suda and H. Shinjoh, Reversible Changes in the Pt Oxidation State and Nanostructure on a Ceria-Based Supported Pt, *J. Catal.*, 2009, **266**, 182–190.
- 155 Y. Nagai, K. Dohmae, Y. Ikeda, N. Takagi, T. Tanabe, N. Hara, G. Guilera, S. Pascarelli, M. A. Newton, O. Kuno, H. Jiang, H. Shinjoh and S. I. Matsumoto, In Situ Redispersion of Platinum Autoexhaust Catalysts: An On-Line Approach to Increasing Catalyst Lifetimes?, *Angew. Chem., Int. Ed.*, 2008, **47**, 9303–9306.
- 156 Y. Nagai, K. Dohmae, Y. Ikeda, N. Takagi, N. Hara, T. Tanabe, G. Guilera, S. Pascarelli, M. A. Newton, N. Takahashi, H. Shinjoh and S. I. Matsumoto, In situ observation of platinum sintering on ceria-based oxide for autoexhaust catalysts using Turbo-XAS, *Catal. Today*, 2011, **175**, 133–140.
- 157 M. Happel, J. Mysliveček, V. Johánek, F. Dvořák, O. Stetsovyeh, Y. Lykhach, V. Matolín and J. Libuda, Adsorption Sites, Metal Support Interactions, and Oxygen Spillover Identified by Vibrational Spectroscopy of Adsorbed CO: A Reference Study on Pt/Ceria Model Catalysts, *J. Catal.*, 2012, **289**, 118–126.
- 158 D. K. Lambert, Stark effect of adsorbate vibrations, *Solid State Commun.*, 1984, **51**, 297–300.
- 159 F. Meng, S. A. Morin, A. Forticaux and S. Jin, Screw Dislocation Driven Growth of Nanomaterials, *Acc. Chem. Res.*, 2013, **46**, 1616–1626.
- 160 X. Yin, J. Shi, X. Niu, H. Huang and X. Wang, Wedding Cake Growth Mechanism in One-Dimensional and Two-Dimensional Nanostructure Evolution, *Nano Lett.*, 2015, **15**, 7766–7772.

