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Thermolysis of 1,3-dioxin-4-ones: Fast generation of kinetic data using in-line analysis under flow

Thomas Durand,^a Cyril Henry,^a David Bolien,^a David C. Harrowven,^a Sally Bloodworth,^a Xavier Franck^b and Richard J. Whitby*^a

Rapid acquisition of kinetic data is demonstrated with a commercial *meso*-scale flow reactor, using a step-change in flow rate or 'push-out' from the flow line. For thermolysis of 1,3-dioxin-4-ones (**1**), we obtain excellent reproducibility in the activation energies measured from spectroscopic data collected by in-line UV or transmission FT-IR monitoring of the output during the transitional period between two flow rates ($\pm 3 \text{ kJ mol}^{-1}$, $0.7 \text{ kcal mol}^{-1}$). Analysis of multi-component UV and IR data is conducted using an orthogonal projection approach (multivariate curve resolution by alternating least squares) for complex spectra, or by calibration-less integration of non-overlapping peak absorbance. All analysis methods were validated using off-line ^1H NMR analysis, and kinetic parameters obtained using the method of a flow rate step-change were validated against conventional steady-state measurements in which time-series data were acquired across multiple experiments. Thermal transfer and dispersion effects are addressed. The experimental methods described herein are valuable for accelerated reaction study and in process development.

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Introduction

Kinetic studies of reactions conducted under continuous-flow conditions present several key advantages over classical 'batch' experiments, including low consumption of reagents, low production of waste,¹ and convenient replication of experiments. In addition, rapid collection of data may be facilitated under flow conditions by in-line analysis, and examples employing UV,² IR,³ Raman,⁴ Fluorescence⁵ and NMR spectroscopy,⁶ as well as mass spectrometry⁷ and HPLC,⁸ have been reported for kinetic study of reactions conducted under steady-state continuous flow conditions. A drawback of using steady-state flow conditions for kinetic studies is that, except where it is feasible to vary the position of an analytical probe along the reactor pathway,⁹ a separate experiment is needed for each time point. An elegant method to overcome this drawback was introduced by Mozharov et al.,¹⁰ whereby collection of time-series data using in-line (Raman spectroscopic) monitoring of the output from a *single* flow experiment is enabled by introducing a step-change in flow rate. A variant upon this method was recently reported,¹¹ in which a flow rate ramp was used to access multiple residence times within a single microreactor flow experiment (with in-line ATR-FTIR monitoring), rather than a step-change in flow rate between two steady states. A purpose-built system for

uninterrupted flow-rate adjustment is required however, and the ramped method is therefore not possible with most current commercial flow platforms.

The aim of our work was to demonstrate that widely available commercial flow reactors could be used for the rapid acquisition of kinetic data for reaction study and process development, enabling wider adoption of the stepped-rate method of 'push-out' from the flow line. The thermolysis of 1,3-dioxin-4-ones (**1**) was used to develop and illustrate the methods developed. We report the use of low-cost UV and transmission FTIR for in-line analysis, address the impact of dispersion (which could be reasonably neglected in the microflow reactor channels used in previous studies), and show that calibration-less multi-component analysis of the resulting spectra provides kinetic data with minimal operator intervention. The results are validated using both off-line ^1H NMR analysis and against known kinetic parameters.

Results and Discussion

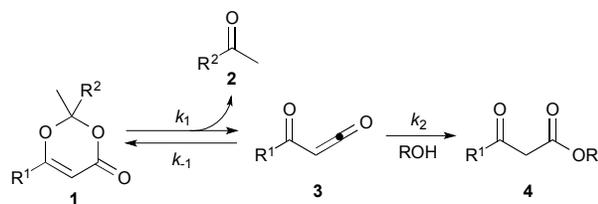
Thermal decomposition of 1,3-dioxin-4-ones **1** via [4+2] retro-cycloaddition,¹² has been well-studied since trapping of intermediate acylketenes **3** with nucleophiles¹³ constitutes a synthetically useful acetoacetylation ($\text{R}^1 = \text{Me}$), for example in the preparation of β -ketoesters **4** when the trapping nucleophile is an alcohol (Scheme 1).

In particular, thermolysis of 2,2,6-trimethyl-4*H*-1,3-dioxin-4-one (**1a**, $\text{R}^1 = \text{R}^2 = \text{Me}$) in the presence of excess alcohol has been shown to be a first-order process¹⁴ in which the rate-determining pre-equilibrium of **1a** and the acyl ketene (**3a**, $\text{R}^1 =$

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Scheme 1 Thermal cycloreversion of 1,3-dioxin-4-ones and acylketene trapping.

Me) is followed by rapid trapping of **3a** by an alcohol ($k_2[\text{ROH}] \gg k_{-1}[\text{acetone}]$)^{13a, 15} such that the kinetic expression reduces as shown in Equation 1. Given the integrated rate law (Eq. 2), we first verified that the rate of reverse trapping of **3** with acetone (k_{-1}) is not significant under flow conditions, by showing that the measured rate constant k does not vary with alcohol stoichiometry. Classical steady-state measurement (using off-line ¹H NMR analysis) gave a consistent first-order rate constant of $k = 1.85 \times 10^{-3} \text{ sec}^{-1} \pm 1\%$ for thermolysis of **1a** in acetonitrile at 120 °C in the presence of 1, 2 or 4 equiv. EtOH.[†]

$$\frac{d[\mathbf{1a}]}{dt} = \frac{k_1 k_2 [\mathbf{1a}] [\text{ROH}]}{k_{-1} [\text{acetone}] + k_2 [\text{ROH}]}$$

$$\frac{d[\mathbf{1a}]}{dt} = k_1 [\mathbf{1a}] \text{ if } k_2 [\text{ROH}] \gg k_{-1} [\text{acetone}] \quad (1)$$

$$\ln[\mathbf{1a}]_t = -k_1 t_r + \ln[\mathbf{1a}]_0 \quad (2)$$

We next sought to apply the push-out method for the study of reaction kinetics under mesoscale flow conditions, using the thermal cycloreversion of **1a** in the presence of 4 equiv. EtOH (premixed) as our proving ground. An excess of the alcohol trap for intermediate acylketene **3a** ($R^1 = \text{Me}$) was used in order to suppress possible side reactions of **3a**; either condensation with acetone (**2a**, $R^2 = \text{Me}$, enol form) to give 2,6-dimethyl-4H-pyran-4-one,^{14, 16} or [4+2] self-cycloaddition to give dehydro acetic acid.^{13a}

Our experiments under flow conditions were conducted using a Vapourtec R series (R2+/R4) system with 10 mL capacity stainless steel reactor of 1mm internal diameter heated by an external oil bath (see below), flow-rate switching within a range of 0.02 – 10 mL min⁻¹, and in-line IR or UV spectroscopic data collection. An alpha transmission FT-IR from Bruker was integrated into the flow system using a Harrick DLC2™ demountable liquid flow cell with sodium chloride windows and IR spectra were recorded with a 3.75 second interval. In-line UV measurements were recorded on an Ocean Optics DH-2000-BAL spectrometer integrated into the flow with a type

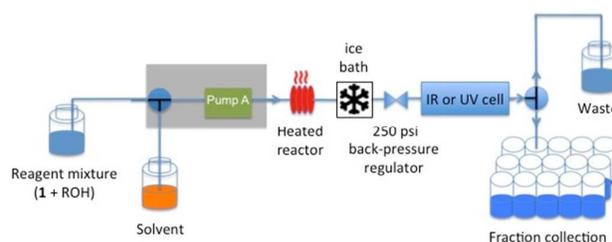


Figure 1 Cartoon of flow set-up utilising a single channel (Pump A) of the Vapourtec R series platform and bottle-feed configuration, 10 mL stainless steel reactor submerged in a precision oil bath, ice bath cooling of the reactor effluent, 250 psi back-pressure regulation and in-line IR or UV data acquisition.

583-F Starna® fluorimeter flow cell (1 mm path length, 0.011 mL volume), enabling collection of UV spectra with a 1.2 second interval. The ability to switch either the described in-line UV or IR analysis into the reactor output flow line means that spectroscopic analysis across a reagent concentration range of 0.001 – 0.3 M is readily achieved. The system was further equipped with an automated sample collector for off-line product analysis (Figure 1).

The method of kinetic data acquisition is described in Figure 2; a typical concentration profile for [**1a**] obtained, in this example, through UV monitoring of the flow output is depicted in Figure 2(a). At each given temperature, once the system has reached steady-state at low flow-rate ($F_1 = \text{typically } 1 \text{ mL min}^{-1}$) a step-change to high flow-rate ($F_2 = \text{typically } 10 \text{ mL min}^{-1}$) is applied, resulting in a second steady-state of correspondingly lower 1,3-dioxin-4-one reaction. The transitional period between steady-states gives spectroscopic information across a gradient of residence times within the reactor. The reaction time (t_r) of individual data points is determined from the experimental time (τ) and inflection points τ_1 and τ_0 (Eq. 3)¹⁰ with further correction for thermal expansion of solvent.¹⁷

$$t_r = \frac{F_2 - F_1}{F_1} \tau + \frac{F_2}{F_1} \tau_1 - \tau_0 \quad (3)$$

Chemometric methods for (calibration-less) spectral deconvolution of multivariate spectra are well developed for quantitative analysis in process monitoring¹⁸ and, in order to reconstruct reactant concentration, multivariate curve resolution of spectra by alternating least squares method (MCR-ALS)¹⁹ was performed using the orthogonal projection approach (OPA).²⁰ In all cases, MCR-ALS resolved two spectral components consistent with clean thermolysis and trapping of **1a**,[†] with the component whose concentration profile correlates with that expected for **1a** displaying a UV spectrum consistent with the 1,3-dioxin-4-one ($\lambda_{\text{max}} 245.9 \text{ nm}$). Representative kinetic data is shown in Figure 2(c) and an Arrhenius plot of the first-order rate constants against the reciprocal of temperature

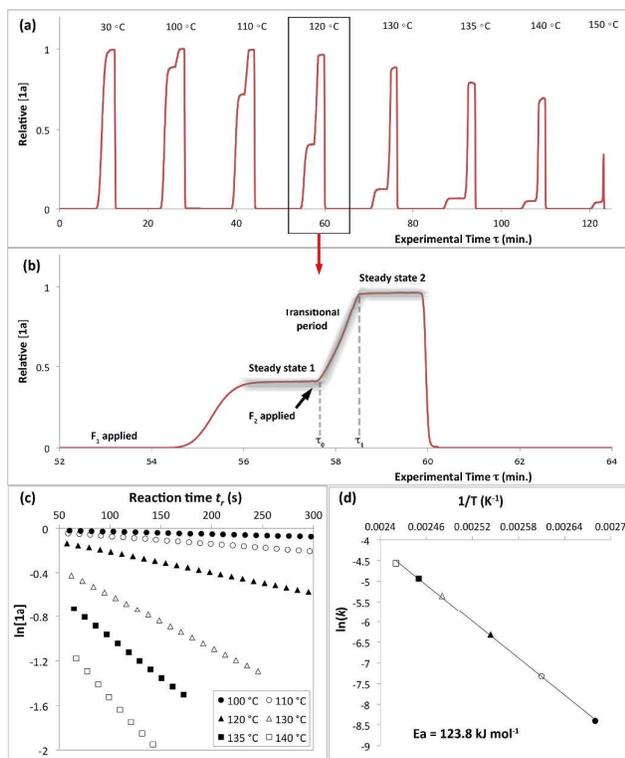


Figure 2 Representative data derived from a push-out experiment (this example is Entry 1 of Table 1, *vide infra*). (a) Concentration profile of 2,2,6-trimethyl-4H-1,3-dioxin-4-one (**1a**) derived from OPA resolution of UV spectra obtained *in-situ* from flow thermolysis in the presence of EtOH (4 equiv.) under stepped flow-rate (push-out) conditions at each of the temperatures indicated. (b) Detailed push-out profile for flow thermolysis of **1a** at 120 °C. (c) First-order kinetics and (d) Arrhenius plots derived from relative [**1a**] data.

[Fig. 2(d)] enables calculation of the activation energy of 1,3-dioxin-4-one thermolysis. Commercially available flow systems are generally designed to work on the *meso*-scale with typical flow rates of between 0.1 and 10 mL min⁻¹, and tube diameters of ~1 mm. Under these conditions flow will be laminar (Reynolds number < 2000) and significant Taylor dispersion²¹ is to be expected. The effect of dispersion is that a sample plug exiting the reactor tubing at a particular time contains material having experienced a range of residence times. The form of the dispersion will have the same residence time distribution (RTD) as that measured for a very short plug of material introduced to and traversing the tube, upon its exit from the tube.

We needed to determine the likely maximum error which dispersion might cause in kinetic studies under push-out conditions. The situation is complicated since the RTD of the material leaving the reactor column at the moment when the step-change in flow rate is applied, is different to the RTD of material which leaves the reactor at the end of the push-out period (because the flow rates while passing through the column are different). Figure 3 shows the observed dispersion of a 10 μL sample of a solution of benzophenone in MeCN, passed through a coiled tubular column of 10 mL capacity and 1 mm internal diameter at flow rates of 10, 1 and 0.2 mL min⁻¹

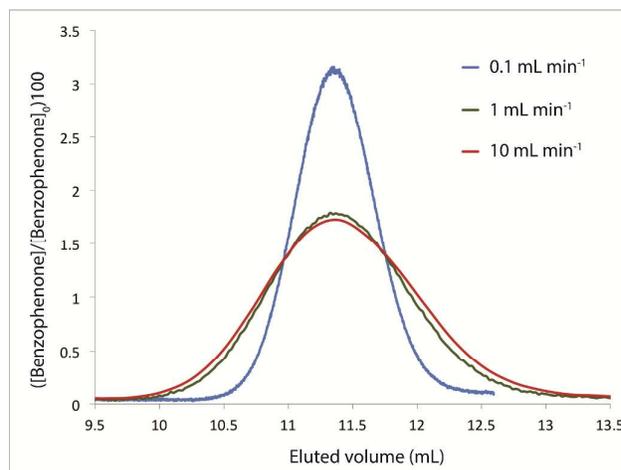


Figure 3 Dispersion of a 10 μL sample plug of benzophenone in MeCN ($[\text{Benzophenone}]_0 = 1.015 \times 10^{-2} \text{ M}$) through coiled stainless steel tubing of 1 mm internal diameter and 10 mL capacity at 30 °C and the flow rates indicated. Concentrations were obtained by UV absorbance.

at 30 °C. The dispersion is shown with respect to volume passing the observation point, rather than time. Dispersion is considerably reduced at higher temperatures and slower flow rates (see supporting information[†]) such that the 10 mL min⁻¹ case in Figure 3 represents a worst case for our studies. It can be seen that the dispersion is significant with respect to the total passage of volume/time, for example the 90% peak width at 10 mL min⁻¹ is around 0.2 min. *cf.* a total spectroscopic sampling time of 1 min. Fortunately the dispersion at higher flow rates is much less than would be predicted from simple Taylor dispersion in a linear capillary due to transverse mixing induced by Dean circulation^{22, 23} as a result of the curved nature of the reactor coil. For the two flow rates used in most of the push-out experiments described below (10 and 1 mL min⁻¹) dispersion is very similar.

Provided that the RTD is symmetric, the average reaction time will be the same as the observed reaction time (*i.e.* that calculated from the residence time according to Mozharov et al.¹⁰). For dispersion profiles obtained at flow rates between 0.1 and 10 mL min⁻¹ at both 30 °C and 120 °C, the difference between the weighted average retention time, and the retention time of the peak top was insignificant. The average conversion will thus be very close to that expected at this time point unless the conversion *vs.* time is highly non linear. Using the dispersion profile from the ‘worst case’ scenario (10 mL min⁻¹, 30 °C) we simulated the effect of dispersion upon calculated kinetic parameters of 1st and 2nd order reactions under push-out conditions, choosing theoretical rate constants corresponding to 80% conversion at the longest reaction time. The effect of dispersion was found to be negligible in each case.[†] An effect is only seen at either end of the push-out period where spectroscopic data under each of the two steady-state flow rate conditions (*e.g.* 1 mL min⁻¹ and 10 mL min⁻¹) is averaged-in; apparent as a slight deviation at each end of a plot of [reagent] *vs.* experimental time (Figure 4) and also by the very slight

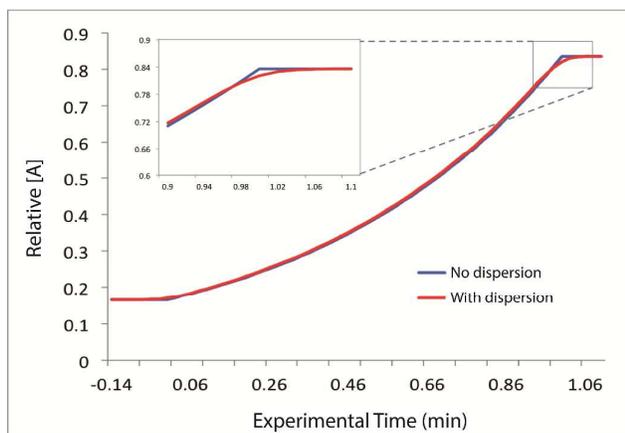


Figure 4 Simulated reagent concentration profiles for a 1st order reaction $A \rightarrow P$ under push-out conditions of stepped flow-rate from 1 mL min^{-1} to 10 mL min^{-1} with flow rate switching at experimental time $t = 0$. Since flow rate increases across the push-out period, residence (reaction) time decreases with increasing experimental time. The indicated profile 'with dispersion' (red line) was generated by redistribution of each data point in the profile with 'no dispersion' (blue line) according to a dispersion profile of benzophenone in MeCN obtained experimentally at $30 \text{ }^\circ\text{C}$ and flow rate = 10 mL min^{-1} in reactor tubing of 10 mL capacity and 1 mm internal diameter.

curvature at each end of the corresponding kinetic plots.[†] Discarding the first and last 10% of the data removes this dispersion effect, although in practice the effect of inclusion on the measured rate constants was negligible.

Thermal transfer was found to be inefficient using the air-heated flow reactor with which the Vapourtec platform is equipped. There was a discrepancy between the steady-state and push-out rate constants at high flow rates which we ascribe to slight cooling of the reactor caused by heat transfer from the air not being sufficient to warm the cold introduced solvent quickly enough in the former experiments. In addition, uneven heating across the double-coil reactor (under control of a single thermocouple) was evident from a kinetic plot of $\ln[1a]$ vs. reaction time, measured for thermolysis of **1a** in acetonitrile at $130 \text{ }^\circ\text{C}$ under push-out conditions, in which two distinct slopes were apparent, corresponding to different measured rate constants in the outer and inner coil. Upon switching to oil-bath immersion of the reactor coil, thermal homogeneity was achieved for kinetic studies. Rate constants of consistent value within standard deviation were measured from linear first-order plots under both steady-state conditions (multiple experiments, each of fixed flow rate between 2 and 10 mL min^{-1} , were used to access multiple residence times in the reactor) and using the push-out method. Data from these heat transfer experiments are

Table 1 Rate constants and activation energies for thermolysis of 2,2,6-trimethyl-4H-1,3-dioxin-4-one (**1a**) in the presence of an alcohol ROH (4 equiv.) in acetonitrile.

Entry ^a	Kinetic Method ^b	Analysis Method	Process Method ^c	F ₁ mL min ⁻¹	F ₂ mL min ⁻¹	ROH	$k \times 10^{-3} \text{ (s}^{-1}\text{)}$								E _a kJ mol ⁻¹
							100	110	115	120	125	130	135	140	
1	Push-out	In-line UV	OPA	1	10	EtOH	0.23	0.66	-	1.81	-	4.70	7.17	10.44	123.8
2	Push-out	In-line UV	Peak abs.	1	10	EtOH	0.23	0.65	-	1.82	-	4.74	7.07	10.81	123.8
3	Steady-state	In-line UV	OPA	-	-	EtOH	0.24	0.68	-	1.89	-	4.92	-	-	126.7
4	Push-out	In-line IR	OPA	1	10	EtOH	0.27	0.71	1.17	1.92	3.16	5.06	7.97	11.58	123.5
5	Push-out	In-line IR	Peak abs.	1	10	EtOH	0.25	0.66	1.09	1.80	3.04	4.78	7.56	11.36	125.0
6	Steady-state	In-line IR	OPA	-	-	EtOH	-	0.72	-	1.83	3.15	5.11	-	11.96	124.4
7	Reverse push-out	In-line UV	OPA	10	1	EtOH	-	0.62	0.96	1.66	2.80	4.47	-	-	128.7
8	Push-out	In-line IR	OPA	1	2	EtOH	-	0.66	1.26	1.96	2.93	4.80	-	-	123.9
9	Push-out	In-line UV	OPA	1	2	EtOH	-	-	1.03	1.74	2.89	4.82	-	11.30	128.2
10	Steady-state	Off-line NMR	-	-	-	EtOH	-	0.65	-	1.89	2.97	5.13	-	12.52	129.9
11	Steady-state	Off-line NMR	-	-	-	BnOH	-	0.67	-	1.79	2.85	4.88	7.74	-	127.9

^a For entries 3, 6, 7, 10 and 11 rate constants were determined from an individual flow experiment; for 'paired' entries 1 + 2, 4 + 5 and 8 + 9 IR and UV data were collected from the same flow output. ^b 'Push-out' refers to the method of stepped flow-rate from F₁ to F₂ with time-series data acquisition across the transitional period within a single experiment, 'steady-state' refers to acquisition of time-series data across multiple experiments (typically five) in which a range of flow rates (each being constant within a given experiment) corresponds to a range of reaction times calculated from the reactor residence time upon correction for thermal expansion of solvent. ^c 'OPA' refers to orthogonal projection approach MCR-ALS analysis, 'Peak abs.' refers to calibration-less calculation of **1a** by integration of non-overlapping peaks in the UV or IR spectrum with respect to 0% conversion. Off-line analysis of ¹H NMR spectra was conducted manually.[†]

given in supporting information.[†]

The results of kinetic studies upon thermolysis of 2,2,6-trimethyl-4*H*-1,3-dioxin-4-one (**1a**) in the presence of 4 equiv. of trapping alcohol in acetonitrile are given in Table 1. Good consistency was obtained between the push-out method and conventional steady-state measurement of kinetic parameters in which time-series data was acquired across multiple experiments (entry 1 vs. 3 and 4 vs. 6). In addition, corroboration of results obtained through calibration-less processing using OPA (MCR-ALS) and using peak absorbance was good (entries 1 vs. 2 and 4 vs. 5). Relative integration of both IR and UV peaks with respect to a $[\mathbf{1a}]_0$ value determined by conducting flow throughput at 25 °C was possible since non-overlapping spectroscopic data were obtained in each case. There was good consistency between the results obtained using UV and IR monitoring (entries 1-3 vs. 4-6), and both were in agreement with the results obtained from manual calculation of $[\mathbf{1a}]/[\mathbf{1a}]_0$ using off-line ¹H NMR analysis (entry 10). The later was repeated using BnOH (entry 11) instead of EtOH as the trap due to concerns about the volatility of the products, with similar results. Detailed data characterisation is given in supporting information although it is interesting to specifically note that, in accord with the clean resolution of mixed component UV spectra mentioned above, no evidence of competing condensation of acylketene intermediate **3a** with acetone (**2a**), or of [4+2] self-cycloaddition of **3a**, was apparent in the NMR spectra.

Excellent reproducibility in the calculated activation energy of cycloreversion of **1a** was achieved overall, $E_a = 126.7 \pm 3.2$ kJ mol⁻¹ (cf. $E_a = 130.1 \pm 4.6$ kJ mol⁻¹ reported by Witzeman and Clemens¹⁴ for thermolysis of **1a** in xylene), independent of the trapping nucleophile (compare entries 10 and 11) as expected. Consistent rate constants were determined regardless of the analysis or process method used and no loss of reproducibility in the push-out experiments was observed across the temperature range studied, suggesting that heat transfer from the heated oil bath to the reactor coil is efficient across this range, such that there is not significant cooling of the initial section of the reactor by entering solvent, even at the high flow rate F_2 .

The short data acquisition time associated with the method of a step-change in flow rate has attracted the criticism that non-instantaneous flow rate adjustment limits the number of data points which can be acquired.¹¹ The use of a ramped change in flow rate to address this issue is not available using most commercial flow systems. Fortunately a simple solution is to be found in a 'reverse push out' method where a step-change from high to low flow rate effects access to time-series data within an (extended) transitional period of increasing reaction conversion (Figure 5). Mozharov et al.¹⁰ have shown that the error associated with non-instantaneous flow-rate adjustment, can be minimised by reducing $\Delta\tau(F_2/F_1)$, where $\Delta\tau$ is the interval between consecutive spectroscopic measurements taken at the reactor output. Under our reverse push-out conditions of flow rate switching from 10 mL min⁻¹ → 1 mL min⁻¹, $\Delta\tau(F_2/F_1) = 0.12$ (UV) or 0.375 (IR). In comparison, 'standard' push-out

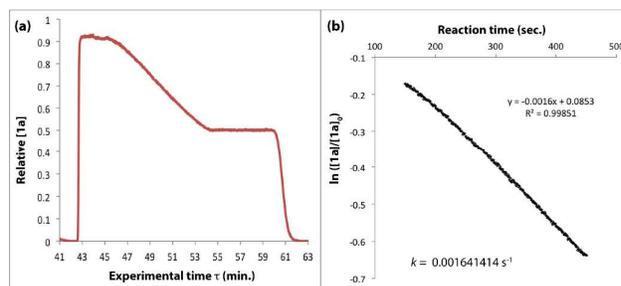
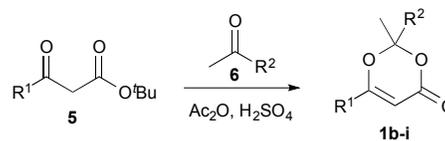


Figure 5 (a) Reagent (**1a**) concentration profile and (b) kinetic plot of 314 datapoints; obtained from OPA resolution of UV spectra recorded *in-situ* for flow thermolysis of 2,2,6-trimethyl-4*H*-1,3-dioxin-4-one in the presence of EtOH (4 equiv.) under reverse push-out conditions ($F_1 = 10$ mL min⁻¹ and $F_2 = 1$ mL min⁻¹) at 120 °C.

(i.e. flow rate switching of 1 mL min⁻¹ → 10 mL min⁻¹) involves a transitional period of ca. 50 sec. and correspondingly $\Delta\tau(F_2/F_1) = 12$ (UV) or 37.5 (IR). Table 1 entry 7 demonstrates that the reverse push-out method gives the same result as the normal push-out method. Whilst greater consumption of reagent is required for equilibration of the flow platform to steady state at a high flow rate, the technique is useful when sampling rate is a limitation. An alternative approach is to use a smaller stepped rate-change of 1 mL min⁻¹ → 2 mL min⁻¹ equating to $\Delta\tau(F_2/F_1) = 2.4$ (UV) or 7.5 (IR). Remarkably good kinetic results are obtained (Table 1, entries 8 and 9), but following a reaction through such a small conversion range is inadvisable since, for example, the linearity of the kinetic plots can no longer be taken to confirm 1st order behaviour.

In order to exploit the fast generation of kinetic data for reaction study, we next sought to conduct a short exploration of substituent effects upon 1,3-dioxin-4-one thermolysis. A series of 1,3-dioxin-4-ones **1b-i**, variably substituted ($R^1 \neq R^2 =$ Me, aryl) at the 2- and 6-positions, were prepared from condensation between a β -substituted- β -keto *tert*-butyl ester²⁴ **5** and ketone **6** under acidic conditions (Scheme 2).²⁵ Activation energies for thermal cycloreversion of **1b-i** were then determined, in the presence of 4 equiv. EtOH in acetonitrile, according to the methods already described (Table 2).



	R ¹	R ²
1b	Me	Ph
1c	Me	<i>p</i> -MeOC ₆ H ₄
1d	Me	<i>p</i> -BrC ₆ H ₄
1e	Ph	Me
1f	<i>p</i> -MeOC ₆ H ₄	Me
1g	<i>p</i> -NO ₂ C ₆ H ₄	Me
1h	2-furyl	Me
1i	2-naphthyl	Me

Scheme 2 Dioxinone synthesis.

For thermal cycloreversion of **1b**, kinetic parameters obtained using the push-out method were again substantiated by close correlation with those obtained from steady-state experiments (compare entries 1 and 3). For both **1b** (entries 1 – 4) and **1c** (entries 5 – 6), different methods of spectroscopic analysis and/or data processing were also found to give consistent values of activation energy. Retro-cycloaddition of **1d-i** was therefore conducted solely using the push-out method with in-line UV analysis and OPA processing (entries 7 and 10 – 14; further scrutiny of the thermolysis of **1d** reported in entries 8 – 9 was conducted subsequently *vide infra*). First-order rate constants for decomposition of **1a-i** gave excellent linear Arrhenius plots (Figure 6) allowing activation energies to be estimated. In general, the differences between activation energies obtained for **1a-i** are too small for detailed interpretation. Nonetheless, the substantial decrease in activation energy observed when $R^2 = p\text{-MeOC}_6\text{H}_4$, together with a significant increase when $R^1 = p\text{-MeOC}_6\text{H}_4$, suggests a transition state in which the C2-O1 bond is substantially broken

(Figure 7). Moreover, the observation that, in several cases, similar rates are associated with different activation energies and *vice versa*, indicates that entropic effects are important. These are likely to be due to the reorganisation of solvent molecules around the sites of developing charge.

The use of OPA or related curve analysis methods is a substantial advance over following the intensity of single peaks in spectroscopic data, particularly in the case of UV monitoring where compounds are characterised by a few very broad, featureless peaks. In the case of **1b, c** and **d** we were concerned that the strength of the UV absorptions of the aryl ketone **2** (Scheme 1) eliminated in the reaction, and their very substantial overlap with the absorption of the reagent 1,3-dioxin-4-ones, may cause difficulty. For these examples we also acquired the kinetic information using IR monitoring, with close results confirming the ability of the OPA method to distinguish between similar chromophores. For **1b** we also confirmed the results using steady-state experiments with both UV and off-line NMR monitoring.

Table 2 Rate constants and activation energies for thermolysis of substituted 1,3-dioxin-4-ones (**1b-i**) in the presence of EtOH (4 equiv.) in acetonitrile.

Entry ^a	1,3-dioxin-4-one	Kinetic Method ^b	Analysis Method	$k \times 10^{-3} \text{ (s}^{-1}\text{)}$										E_a kJ mol ⁻¹
				85	90	95	100	105	110	115	120	125	130	
1	1b	Push-out	In-line UV	-	0.24	0.45	0.76	1.32	2.19	3.54	5.75	-	-	125.2
2	1b	Push-out	In-line IR	-	-	0.48	0.74	1.27	2.08	3.55	5.84	-	-	121.4
3	1b	Steady-state	In-line UV	-	-	-	0.80	1.31	2.30	3.70	6.16	-	-	124.9
4 ^c	1b	Steady-state	Off-line NMR	-	-	-	0.79	1.24	2.33	3.29	6.26	-	-	124.6
5	1c	Push-out	In-line UV	1.66	2.78	4.71	7.55	11.49	17.18	-	-	-	-	107.2
6	1c	Push-out	In-line IR	-	2.65	4.52	7.15	11.11	-	-	-	-	-	108.6
7	1d	Push-out	In-line UV	-	-	-	0.59	0.94	1.59	2.68	4.54	7.80	-	128.3
8	1d	Push-out	In-line IR	-	-	-	0.56	1.06	1.82	2.91	-	7.47	-	126.6
9	1d	Push-out	Off-line NMR	-	-	-	0.55	0.95	1.62	2.60	-	7.42	-	128.0
10	1e	Push-out	In-line UV	-	-	-	-	0.32	0.55	0.93	1.56	2.58	4.04	128.6
11	1f	Push-out	In-line UV	-	-	-	-	-	0.44	0.76	1.30	-	3.62	136.0
12	1g	Push-out	In-line UV	-	-	-	0.57	0.91	1.53	2.47	4.15	-	-	120.9
13	1h	Push-out	In-line UV	-	0.37	0.61	1.05	1.72	2.99	4.95	-	-	-	122.3
14	1i	Push-out	In-line UV	-	-	-	0.21	0.34	0.59	0.96	1.60	-	3.98	124.0

^a Each entry quotes rate constants determined from an individual flow experiment. ^b 'Push-out' refers to the method of stepped flow-rate from $F_1 = 1 \text{ mL min}^{-1}$ to $F_2 = 10 \text{ mL min}^{-1}$ with time-series data acquisition across the transitional period within a single experiment. 'Steady-state' refers to acquisition of time-series data across multiple experiments (typically five) in which a range of flow rates (each being constant within a given experiment) corresponds to a range of reaction times calculated from the reactor residence time upon correction for thermal expansion of solvent. All data was processed using orthogonal projection approach MCR-ALS analysis with the exception of analysis of ¹H NMR spectra (entries 4 and 9) which was conducted manually. ^c BnOH (4 equiv.) was used as the alcohol trap.

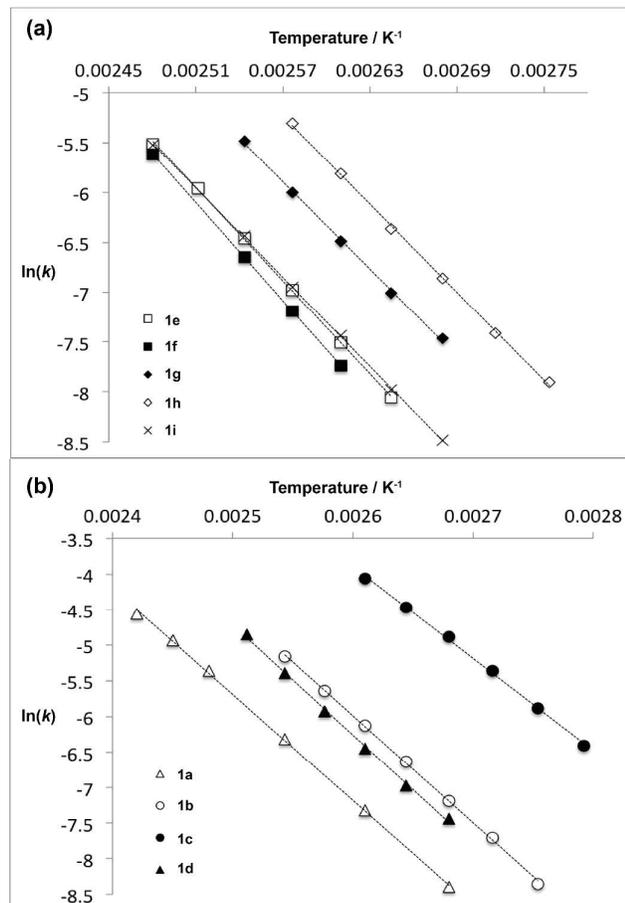


Figure 6 Arrhenius plots for thermal decomposition of (a) 1,3-dioxin-4-ones **1e-i** with different 6-position substituent R^1 and (b) 1,3-dioxin-4-ones **1a-d** with different 2-position substituent R^2 .

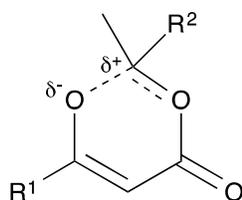


Figure 7 Dipolar transition state.

For compound **1d** we also introduced the use of NMR to monitor a push-out experiment by using a fraction collector to split the output into 20 samples which were then analysed by off-line ^1H NMR. The resulting E_a (Table 2 entry 9) was in excellent agreement with that obtained using in-line UV and IR analysis (*cf.* entries 7 and 8).

Conclusions

The ‘push-out’ technique using a stepped flow rate in a continuous flow reactor has been developed using a widely available commercial flow reactor, to rapidly obtain kinetic data of suitable quality for application in process development chemistry.

Integration of in-line UV and transmission FT-IR provides a versatile breadth of spectroscopic methods for data collection, suitable for multiple reaction types. We have also demonstrated the use of orthogonal projection approach MCR-ALS for deconstruction of multi-component spectroscopic data collected in-line under push-out conditions, and validated these methods against off-line ^1H NMR data collection with traditional manual analysis. We have shown that an extended period of data collection is readily accessible using a ‘reverse’ push-out of the flow line upon switching from high to low flow rate.

Overall, the methods presented offer great potential for wider adoption of the push-out method to obtain kinetic data, providing opportunities for accelerated reaction study and process development.

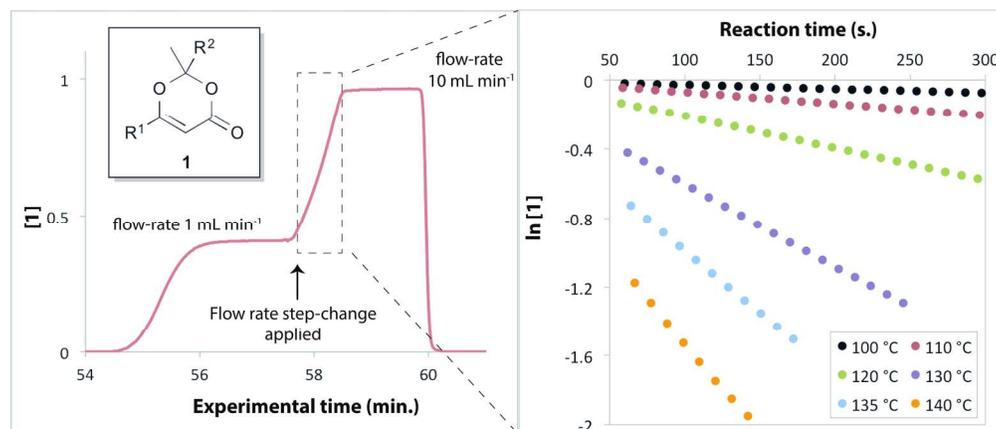
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Rapid acquisition of kinetic data for thermolysis of 1,3-dioxin-4-ones is demonstrated with a commercial meso-scale flow reactor, using a step-change in flow rate or 'push-out' from the flow line.
175x75mm (300 x 300 DPI)