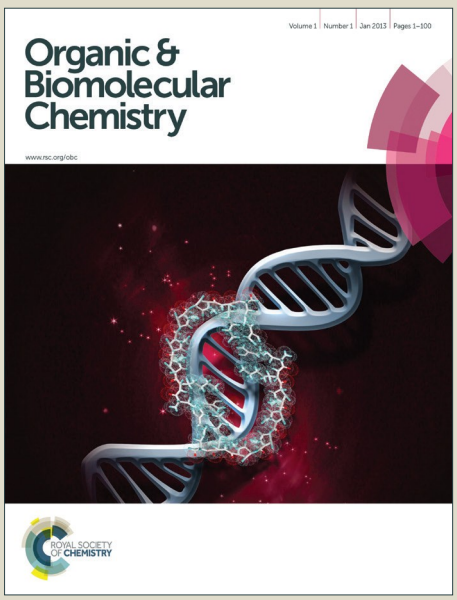
Organic & Biomolecular Chemistry

Accepted Manuscript



ROYAL SOCIETY OF CHEMISTRY

www.rsc.org/obc

Multi-conformer molecules in solutions: an NMR-based DFT/MP2 conformational study of two glucopyranosides of vitamin E model compound†

Ryszard B. Nazarski,‡*[a] Piotr Wałejko[b] and Stanisław Witkowski[b]

Abstract: Overall conformations of both anomeric per-*O*-acetylated glucosyl derivatives of 2,2,5,7,8-pentamethylchroman-6-ol were studied in the context of their high flexibility, on the basis of NMR spectra in $CDCl_3$ solution and related DFT calculation results. A few computational protocols were used including diverse density functional/basis sets combinations with special emphasis on accounting (at various steps of the study) for an impact of intramolecular London-dispersion (LD) effects on geometries and relative Gibbs free energies (ΔGs) of different conformers coexisting in solution. The solvent effect was simulated by an IEF-PCM approach with the UFF radii; its other variants, including the use of recently introduced IDSCRF radii, were employed for a few compact B3LYP-GD3BJ optimized structures showing one small imaginary vibrational frequency. The advantage of IDSCRF radii for such purposes was shown. Of the four tested DFT methods, only application of the B3LYP/6-31+G(*d*,*p*) approximation afforded ensembles of 7-8 single forms for which population-average values of computed NMR parameters (δ_H, δ_C and some $^nJ_{HH}$ data) were in close agreement with those measured experimentally; binuclear ($\delta_{H,C}$ 1:1) correlations, $r_{H,C}^2 = 0.9998$. The associated individual ΔG values, corrected for LD interactions by applying the Grimme's DFT-D3 terms, afforded relative contents of different contributors to analyzed conformational families in much better agreement with pertinent DFT/NMR-derived populations (*i.e.,* both data sets were found practically equal within the limits of estimated errors) than those calculated from dispersion uncorrected ΔGs. All these main findings were confirmed by additional results obtained at the MP2 level of theory. Various other aspects of the study such as crystal *vs.* solution structure, *gg*/*gt* rotamer ratio, diagnostic (de)shielding effects, dihydrogen C-H···H-C contacts, and doubtful applicability of some specialized DFT functionals (M06-2X, ωB97X-D and B3LYP-GD3BJ) for description of highly flexible molecules are also discussed in detail.

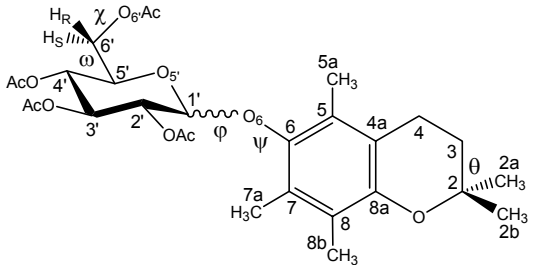
Introduction

High-resolution nuclear magnetic resonance (NMR) spectroscopy is undoubtedly the most valuable experimental technique for determination of structure and dynamics of small- to medium-sized organic molecules, especially carbo- and heterocyclic ones, when elucidating a relative configuration and/or assessing the overall multi-conformer (composite) geometries[1] of such species in solution. Among various isotropic NMR parameters, chemical shifts, δ_Ks, and indirect spin-spin coupling constants (hereafter J_{KL} couplings) are the most informative observables employed for such purposes. Nowadays, these possibilities have become considerably enhanced for common spin-½ magnetic active nuclei by two methods of NMR-oriented density functional theory (DFT) calculations, *i.e.,* gauge-including atomic orbital (GIAO)[2] predictions of absolute shieldings σ_Ks (and thus interrelated δ_K data) and computations of J_{KL} values.[3] The use of such approaches is crucial for structurally flexible systems affording only population-weighted averaged NMR spectra in solution. Indeed, reliable calculations of the above (not accessible in another way) NMR parameters for the major contributors to their conformational ensembles, are usually necessary in all cases of this kind.[1]

[a] *University of Łódź, Faculty of Chemistry, Department of Theoretical and Structural Chemistry, Pomorska 163/165, 90-236 Łódź, Poland. E-mail: nazarski@uni.lodz.pl; Fax: (+48) 42635-5744; Tel: (+48) 42635-5615*
[b] *University of Białystok, Institute of Chemistry, Ciołkowskiego 1K, 15-245 Białystok, Poland*

†Electronic supplementary information (ESI) available: Experimental – General information, Synthesis and spectroscopic data, NMR spectra of **1α**, **1β** and **2α**, Δ_{H5} values for selected glucosides; Calculation results for **1α** and **1β** – Structural parameters, Energetics and populations, Molecular representations, Scatter plots of J_{HH}^{cald}s *vs.* J_{HH}^{obsd}s and $\delta_{H,C}^{calcd}$(MP2) *vs.* $\delta_{H,C}^{obsd}$ relationships, Cartesian coordinates of DFT structures. See DOI: ……..
‡Physical image *vs.* molecular structure relation, Part 19. For Part 18, see ref. 91c.

2,2,5,7,8-Pentamethylchroman-6-ol (PMC) – the parent system of title compounds – is a potent phenolic free radical scavenger related to α-tocopherol (vitamin E),[4] in which a long lipophilic phytyl side chain is replaced by a methyl (Me) substituent. It is the potent hydrophilic α-tocopherol derivative,[5] but its biological activity is not always shared by its parent α-tocopherol (*e.g.,* it acts as a potent anti-inflammatory agent).[5b] PMC shows over 5-10 times stronger dose-dependent inhibition of the agonist-induced platelet aggregation in human platelet-rich plasma, if compared to α-tocopherol.[6] Among various α-tocopherol analogues, it is the most potent inhibitor of nuclear factor-kappa B (NF-κB) activity.[7] Moreover, PMC has shown anti-androgen activity in prostate carcinoma cells and is considered a potent chemopreventive agent of androgen-dependent diseases, such as prostate cancer[8] and other human cancers.[9] Nevertheless, the potential therapeutic application of PMC is limited due to its relatively low water solubility. One of the most promising solutions is to convert PMC into its amphiphilic glycoconjugates.[10] These derivatives as prodrugs would gain a favorable solubility in physiological fluids and a proper permeability through membranes and natural biological barriers *e.g.,* blood to brain. New data indicate that PMC can be helpful in the designing of such new potential medicinal compounds of a better clinical effectiveness.[11] Some glycosides of vitamin E and its short-chain analogues were described earlier.[10b,12] Also different structural aspects of this type and related model molecules, such as PMC and Trolox, were studied extensively in our laboratory by means of ^{13}C NMR in solution[13] and the solid state (CP/MAS technique)[14] as well as by ECD spectroscopy.[15] It is obvious, that for complete understanding a behavior of every system having potential biomedical activity, a good knowledge of its conformational properties (both structure and dynamics) is crucial. Therefore, a comprehensive ^{1}H and ^{13}C NMR data-based DFT conformational investigation of the two peracetylated glucosyl derivatives of PMC, *i.e.,* compounds **1α** and **1β** (see Fig. 1), was undertaken.



1α (*ax* OAr)
1β (*eq* OAr)

Fig. 1. Structures of the studied compounds with the atom numbering and five relevant torsion angles concerning their mobile molecular units, where Ar means the chroman system.[16]

In view of the foregoing, the two title highly structurally flexible 2,3,4,6-tetra-*O*-acetyl-D-glycopyranosides seemed to be also entities particularly suitable for testing of a few calculational DFT-level protocols currently available for the analysis of composite shapes[1] of small- to medium-sized multi-conformer systems. Indeed, such mobility concerns even the aglycone (non-sugar) moiety of **1α** in the solid state, as its 3,4-dihydro-2*H*-pyranyl (DHP) ring adopting two alternative half-chair (HC) like forms was found disordered in the crystal structure at 100 K, along with related *gem*-dimethyl groups.[18] Hence, ^{1}H and ^{13}C NMR spectra of both anomers of **1** in $CDCl_3$ were fully interpreted and additionally analyzed in the light of σ_H and σ_C values GIAO-predicted for their preselected energetically reasonable forms. Some diagnostic J_{HH} and J_{CH} couplings were calculated as well. The integral equation formalism (IEF)[19] of an implicit solvation and UFF-radii cavities were mainly used within the polarizable continuum model (PCM)[20] approach. Its other variants were also employed for some structures with one small imaginary harmonic vibrational (IHV) frequency, showing an advantage of the use of the recently introduced[21] and applied[22] IDSCRF radii in such

cases. Moreover, an empirical *a posteriori* correction of the computed Gibbs free energy, *G*, data[1c,23] for a proper account of long-range London dispersion (LD) forces of the van der Waals (vdW) type, which are neglected in conventional DFT approaches (with underestimation of LD),[24,25] was *inter alia* tested with the use of pairwise DFT-D3 corrections of Grimme.[25c]

Thus, four inseparable points were specially addressed in this work: (*i*) a good representation of the overall solution shapes[1] of glucopyranosides **1α** and **1β** considered as highly flexible molecules, (*ii*) testing of a few DFT model chemistries (functional and basis set) accessible today for most reliable prediction of the structure and molecular, *e.g.,* NMR spectroscopic, properties of the individual forms of **1** coexisting in real solutions at equilibrium, and, particularly, (*iii*) an *explicit* accounting for the impact of weak intramolecular LD attractions[24,25] on separate geometries and/or (*iv*) a *post factum* accounting for the influence of such interactions on their relative conformational energies (ΔG^{o}). To the best of our knowledge, this kind of a widespread NMR data- and dispersion-oriented DFT study of the multi-conformer systems, positively verified by additional results emerging from the much more expensive MP2 calculations, has not yet been published.

Results and discussion

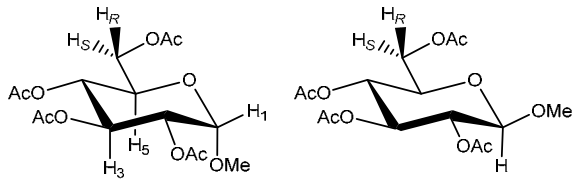
NMR spectra of 1 and other related systems

The title *O*-glucopyranosides were synthesized from PMC[26] according to the reported procedure[10b] based on the Helferich glycosylation method,[27] using peracetylated β-glucose as a donor and a mild Lewis acid ($ZnCl_2$) as a glycosyl promoter, followed by deacetylation.[28] The resulting deprotected α/β-anomers were separated chromatographically and then subjected to acetylation (Experimental†). The isolated products **1α** and **1β** gave spectral data fully consistent with the literature.[10b] The molar α/β ratio of 36:64 (by ^{1}H NMR) was established when pure **1β** was melted with $ZnCl_2$ (1.2 equiv) at 390 ± 5 K under diminished pressure (30 Torr), whereas **1α** was decomposed with the liberation of PMC under the same conditions.

Analysis of various types of NMR spectra recorded for **1α** and **1β** in a $CDCl_3$ solution at a 14.1 T magnetic field strength (for 1D spectra, see Figs. S1-S6†), was performed as previously for the other multi-conformer systems.[1,29] Thus, the δ_H, δ_C and $^3J_{H,H}$ values associated with their anomeric centers were found in agreement with those for D-glucopyranoses.[30] Also all cross peaks due to expected C-H connectivities within both molecules were localized in 2D spectra, including correlations across the glucosidic linkage in ^{1}H,^{13}C HMBC plots. Moreover, diagnostic $^1J_{C1',H1'}$ couplings (of 172.1 and 163.4 Hz for **1α** and **1β**, respectively) fully compatible with literature data (*ca.* 170–175 and 160–165 Hz for α- and β-forms, respectively)[30b] were derived from HMBC spectra. Only assignments of the two slightly differentiated NMR lines coming from protons/carbons in 2a/2b-*gem*-dimethyl groups and two ^{13}C lines of the C3'/C6' acetate methyl groups were not provided by an NMR experiment, however all these signals were unambiguously assigned in further calculations (*vide infra*). An observed chemical shift non-equivalence of these former Me groups indicated that the C6–O6 rotation is not (nearly) a free-energy process, because sharp ^{1}H/^{13}C resonance lines of the 2a/2b geminal groups are observed for PMC and its derivatives.[13a,26b] On the other hand, cross peaks of the four H1'/H5a (where H1' ≡ C1'-H, *etc.*), H1'/H7a, H5'/H5a and H5'/H7a pairs and the first two ones were found in ROESY spectra of **1α** and **1β**, respectively, as arising from related inter-residual H-H contacts. These nuclear Overhauser effect (NOE) data, well corroborated by broadening of a vast majority of the ^{1}H signals of aglycone moieties of both anomers (Figs. S1 and S4†), confirmed a high degree of rotameric flexibility around their C1'–O6 and/or neighboring C6–O6 bonds. In turn, conformational mobility of

the constituent semi-unsaturated DHP rings is additionally indicated by motional averaging of the $^{2}J_{H3,H4}$ values (Experimental†).

Interestingly, two long-range couplings, $^{4}J_{H1',H3'} = 0.4_0$ and $^{4}J_{H1',H5'} = 0.5_2$ Hz, were revealed for a sugar residue of **1α** in ^{1}H NMR spectra processed with resolution enhancement.[31] Similar interactions ($^{4}J_{H1,H5} = 0.5_4$ and $^{4}J_{H1,H3} = 0.3_6$ Hz) were also determined for methyl 2,3,4,6-tetra-*O*-acetyl-α-D-glucopyranoside **2α** as the simplest aliphatic analogue of **1α** (Fig. 2). The latter results are in good agreement with corresponding heteronuclear couplings $J_{C,H}$s found for **2α** in C_6D_6 ($^{3}J_{H1,C5} = 6.9$ and $^{3}J_{H1,C3} = 5.6$ Hz).[32] Moreover small couplings, $^{4}J_{HH}$ ~0.4_5 Hz, between protons of the *gem*-dimethyl groups were found in both title systems. To our best knowledge, all these $^{4}J_{HH}$ couplings, whose existence was confirmed by our calculations (*vide infra*), were not reported before and were possibly unobserved.[33] So, the sharp well-resolved multiplet consisting of 16 lines (dddd) due to an axial H5' appeared in $CDCl_3$ at 4.528 and 3.986 ppm for **1α** and **2α**, respectively (Figs. S2 and S8†). Additionally, clear NOE interactions H4'/H6'*R* were observed in ROESY spectra of both these α-anomers confirming the configuration at C6'.[32] Among other issues, the origin of a 0.54 ppm-variation in the above ^{1}H chemical shifts, and, especially, a pronounced 1.00 ppm-difference $\Delta_{H5'} = \delta_{H5'\alpha} - \delta_{H5'\beta}$ found for compounds **1** was a particularly intriguing question. Such large $\Delta_{H5'}$ values were also determined for anomeric pairs of other *O*-glycosides of chroman-6-ol (**3-8**) possessing *inter alia* the mannoside, galactoside or 2-deoxyglucose residue; for details see Tables S1-S3†. Furthermore, it was found that Δ_{H5} diminished with the change of bulkiness of the aglycone moiety, but an impact of the pyranose ring structure is evident as well – compare the Δ_{H5} value (1.00 *vs.* 0.70 ppm) for **1** and **8**, respectively (Table S3†). Nonetheless, one can conclude that Δ_{H5} is a much better determinant of stereochemistry at C1 than the usually considered difference $\Delta_{C1} = \delta_{C1\alpha} - \delta_{C1\beta}$, at least for glycopyranosides **1-8** (Table S2†).



2α **2β**

Fig. 2. Structures of two anomeric methyl glucosides.

Conformational study

Owing to complexity and great flexibility of both glucosides **1**, their conformational analysis was done on the basis of structural information available from the NMR data, which was supplemented with computational results. Thus, a few standard approaches were applied in two inseparable steps of the study. An extensive HF/DFT modeling of the series of low-energy candidate conformational states of both anomers **1** was performed at beginning, by starting with huge amounts of their molecular-mechanics (MM) models found initially. This step was followed by predictions of relevant NMR spectral parameters (J_{KL} and mainly δ_K values) of such DFT-optimized structures, carried out with the use of different combinations of density functionals and basis sets (Calculational). Moreover, due to a fortunate lack of strong specific solvent-solute interactions, their solvation was simulated within the framework of an implicit solvation model, by using mainly an IEF-PCM[19,20] method as implemented in the Gaussian 09 package of programs.[34] Based on the $G_{DFT\ 390}$ values computed in the standard way, it was found that **1β** more thermochemically stable than **1α**, but agreement with the equilibration experiment mentioned above was only qualitative (see however below).

In order to find the fully relaxed overall shapes[1] of molecules **1α** and **1β** in the most general manner, a linear regression analysis of the measured δ_H and δ_C data *vs.* those values obtained from the σ_H/σ_C data GIAO predicted at the IEF-PCM(UFF,$CHCl_3$)/mPW1PW91/6-311+G(2*d*,*p*)[35]//IEF-PCM(UFF,$CHCl_3$)B3LYP/6-31+G(*d*,*p*)[36] level was performed for some promising forms found at the beginning. The double-ζ (DZ) valence quality of employed atomic basis sets was forced by the relatively large size of the molecules under study. Thus, the calculated data were plotted as usual[1,23b,37] against experimental ones, but using the binuclear $\delta_{H,C}^{calcd}$(DFT) *vs.* $\delta_{H,C}^{obsd}$ correlation;[1a,37,38] see Computational. The so obtained individual NMR data-derived populations p_i were next confronted with pertinent results on energetics of different single forms of **1α** (or **1β**) coexisting in solution at equilibrium, *i.e.,* relative total electronic-nuclear energies (0 K, ΔE_{tot}s) or relative standard Gibbs free energies (298.15 K, ΔG^os), computed for local minima on conformational energy hypersurfaces of analyzed solutes immersed in a polarizable continuum, of which relative permittivity matches that of $CHCl_3$.

The above preliminary calculational *vs.* experimental data sets were subsequently analyzed in light of our previous results on the other non-rigid (flexible) molecules.[1,23b] In particular, the reliability of a standard approach concerning energy-weighted fractional populations[39] and the reproduction of weak long-range attractive intramolecular LD forces of the vdW type,[24,25] operative in two relative large systems **1**, were considered. Thus, all available data were analyzed in terms of Boltzmann populations of potential contributors to the overall composite shapes of both of these molecules, based on the *G* values computed for their individual conformers in simulated solutions. The structure of glycosides is usually described[40] by two torsion angles around the glycosidic linkage, *i.e.,* φ (O5'–C1'–O6–C6) and ψ (C1'–O6–C6–C5), and the ω angle (O5'–C5'–C6'–O6') within the exocyclic acetoxymethyl group (Figs. 1 and 3). Hence, great rotameric flexibility is generally possible, but only some of the above rotamers of **1α** and **1β** really exist in solution. In other words, their conformational freedom was found to be restricted to only a few (nearly) freely rotating bonds, as described later.

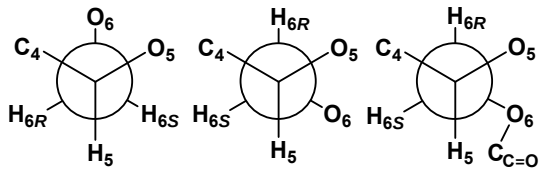


Fig. 3. Newman projections outlining the nomenclature used throughout for the discussed C5'–C6' rotamers.

Fortunately, the first of the angles mentioned above was found at the same magnitude (φ = 127° and *ca.* −73° for **1α** and **1β**, respectively) in all our initial B3LYP-optimized structures, *i.e.,* 8 forms of **1α** and 7 forms of **1β**, derived from the respective starting MMX geometries (Computational). The D-glucopyranose ring of both systems was consistently computed to be a unit adopting the relative rigid $^4C_1$ chair conformation.[41] Also the three consecutive equatorial acetoxy groups in positions 2', 3' and 4' were always found situated in the planes approximately perpendicular to an average sugar plane, in line with such arrangements determined in the crystals of **1α**[18]and **1β**.[42] Moreover, one of the three rotamers (each separated by ~120° dihedral rotation) around the exocyclic C5'–C6' bond in a pyranose ring, *i.e.,* the *tg* form[43] with ω ≈ 180°, was not found within the used 25.1 kJ mol^{-1} MMX energy cutoff. This finding was in agreement with the assumption that little or no contribution would be expected from the *tg* rotamer of **1**, because of unfavorable steric interactions between the acetoxy groups borne by C4' and C6'.[32,43a] Indeed, its participation for anomers **2α** and **2β** having an identical glucose residue was suggested[32,43c] as only 4-11 and 1-8%, respectively, so practically within an

estimated uncertainty of 5-10% in the NMR data-based conformer population.[43c] Thus, the other three staggered rotamers [namely *gg* ($\omega \approx -60^o$),[43] *gt* ($\omega \approx 60^o$, $\chi \approx \pm 180^o$)[43] and an unusual 'bent' form denoted hereafter gt_{90} ($\omega \approx 60^o$, $\chi \approx 90^o$), all shown schematically in Fig. 3 and characterized by the angles ω and χ (≡ C5'–C6'–O6'–$C_{C=O}$) given in parentheses], four O6–C6 rotamers [referred to as $R\alpha^-$ ($\psi \approx -62^o$) and $R\alpha^+$ ($\psi \approx 123^o$) for **1α** as well as $R\beta^-$ ($\psi \approx -80.5^o$) and $R\beta^+$ ($\psi \approx 104.5^o$) for **1β**, with the ψ values stated above] and two half-chairs arising from the ring-puckering deformation of a DHP moiety,[18] *i.e.,* HC^- ($\theta = -58.5^o$) and an opposite form HC^+ ($\theta \approx 58.5^o$) with the angle θ = C1–C2–C3–C4, were analyzed in detail. Hence, the twelve most promising candidate structures with all possible combinations of the local atom arrangements (geometric motifs) mentioned above, which were originally found applying the GMMX random subroutine of PCMODEL[44] (above 15 forms), constructed from incomplete geometries of two crystallographically independent molecules coexisting in the crystal structure of **1α** (2 forms)[18] and additionally built with the MM+ force field[45] of Hyperchem[46] by adequate modification of the geometry of other forms in our hand (7 remaining forms),[47] were taken into account in all further studies for every two molecules (for all details, see Tables S4 and S5†). In both structures found in the crystal of **1α**, the CH_2OAc unit adopts the gt_{90} form. As far as we know, the presence of such 'bent' rotamers in solution was not considered before.

However, the rather highly incoherent conformational landscape was found in a general manner outlined above. Indeed, several trial structures of **1α** and **1β** proposed as privileged on the basis of standard Δ*G* data (and for which all GIAO-based δ_H and δ_C values were *a priori* predicted) were 'not visible' in the measured NMR spectra. More precisely, simulated ^1H and ^{13}C chemical shifts, obtained as Boltzmann-population-weighted sums of such NMR parameters computed for these individual forms of **1α** and **1β**, did not match related values found experimentally. An occasional failure of such a common approach[39] for flexible molecules is poorly documented in the literature dealing with NMR[1,49] and infrared vibrational circular dichroism (VCD)[50] spectroscopic studies in solutions. The usage of a '*solution-phase environment (spectroscopic) match criterion*' instead of an '*energetic criterion*' was suggested in some cases.[1] These discrepancies most likely originate from known imperfections of used theoretical approaches, *e.g.,* not adequate mimicking the influence of surrounding media[1c,49,50] and/or accounting for LD effects[1c] for multi-conformer systems, *i.e.,* geometries, relative energies (Δ*G*s) and spectral responses of single contributors to their conformational families in solution. But we must also keep in mind that in certain physico-chemical and biophysical events wide energy basins associated with ensembles of many structurally similar, highly flexible conformers ('flat' local minima) may be preferred over narrow energy wells of comparable depth and representing individual rigid forms (global minima), owing to the entropy factor.[51]

In view of the foregoing, the B3LYP-GD3BJ[25c,e,34] flavor of DFT corrected for dispersion energy was applied in additional geometry reoptimizations carried out 'in $CHCl_3$' starting with the 24 most promising B3LYP structures discussed earlier. However, all of these computational efforts, performed again using the standard IEF-PCM approach with UFF radii-cavities, led to very disappointing results. Indeed, much worse agreement between values of predicted and observed NMR parameters was generally found for the structures of **1α** and **1β** optimized in this way (data not given). Analogous effects were also obtained with two other specialized DFT functionals, namely, M06-2X[52] and ωB97X-D.[53] Thus, the M06-2X structures were similar to extended B3LYP geometries, while more compact shapes predicted with ωB97X-D were close to those B3LYP-GD3BJ optimized (for views of selected forms, see Figs. 4, S12 and S13†). These new geometries of **1α** and **1β**, described in terms of five torsion angles (Fig. 1), are collected in Tables S4 and S5† together with those of initial B3LYP structures. Also pertinent Gibbs free energies are given there, with the exception of such values for some B3LYP-GD3BJ optimized geometries having one small IHV frequency (up to $10i$ cm^{-1}). For

these latter structures some uncertainties in their G^o values are expected, because the constituent zero-point vibrational energy (ZPVE) term is calculated only from non-imaginary frequencies.[54]

The main modifications of the geometry of **1α** and **1β** concern the angle φ, which increases from *ca.* 125° to 156° for **1α**, and χ which decrease from *ca.* −180° to −123° or even −107° for **1α** and **1β**, respectively. Particularly large rotational freedom, manifested by a relatively wide low-energy valley, exists for the C1'–O6 rotation in **1α** (Tables S4 and S5†). The greatest changes are observed on going from B3LYP to compact structures ωB97X-D and especially B3LYP-GD3BJ (*cf.* Figures S10 and S11 *vs.* S12 and S13†). Thus, large movement of the aromatic part of aglycone to the C2' acetoxy groups primarily takes place for most forms of **1α** (arrangement of the type I, changes in φ), while the C6' acetoxy units in all their *gg* rotamers move strongly towards the C4' acetoxy groups with unexpected formation of *tg* forms *via* a C5'–C6' rotation (type II, χ → *ca.* −123°); the latter displacement is less pronounced for M06-2X (χ → *ca.* −148°). In turn, Me groups of C6' acetoxy units in *gt* rotamers of **1β** move strongly towards 2a-Me group in a DHP ring of aglycone with the formation of stabilizing C–H···H–C attractions[55] (type III, χ → *ca.* −108°); this change is marginal for M06-2X (χ → *ca.* −167°). Two B3LYP-GD3BJ-optimized non-physical geometries of $\mathbf{14599}_{\mathbf{comp}}$ and $\mathbf{12272}_{\mathbf{comp}}$[56] with LD effects of the type I/II and III (these latter exemplified in $\mathbf{12272}_{\mathbf{comp}}$ by two short dihydrogen CH···HC contacts of 2.327 and 2.753 Å), respectively, are shown in Fig. 4. The displacements of the type I are a little similar to stabilizing intramolecular attractions originating in LD forces between aromatic rings and π–electron containing groups found in high-level correlated ab initio calculations concerning some oligopeptides and isolated small proteins in the gas phase.[25a,b,57]

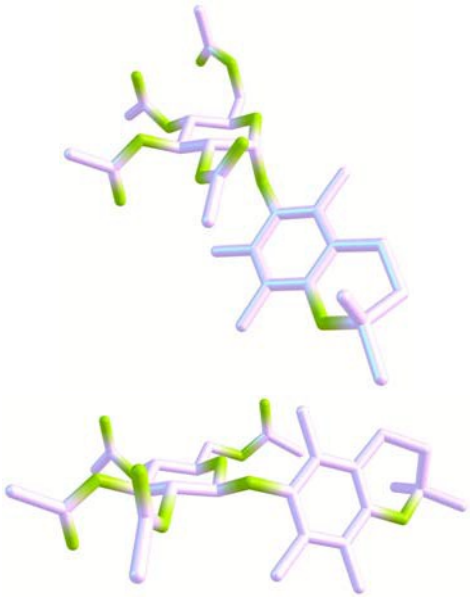


Fig. 4. Chemcraft 3D drawings of two types of non-physical structures found at the IEF-PCM(UFF,$CHCl_3$)/B3LYP-GD3BJ/6-31+G(*d*,*p*) level: $\mathbf{14599}_{\mathbf{comp}}$ (top) and $\mathbf{12272}_{\mathbf{comp}}$ (bottom), see the text for details.

The aforementioned atom displacements and especially the presence of *tg* forms instead of *gg* ones in analyzed solution mixtures (contrary to the observation, *vide supra*) was perhaps a reason for the lack of consistency between computed and measured NMR data. Hence, one can suppose that *use of* these specialized DFT functionals (*ωB97X-D and B3LYP-GD3BJ, in particular*) *is rather not suitable for modeling the ground-state geometry of the title and most likely also other floppy molecules with small barriers to conformational changes* owing to an overestimation of LD attractions. Some recent examples of transition-state structures optimized by these or other similar

methods – and for which also too strong intramolecular LD effects (and so not wholly reliable ΔGs) were computed – were reported for B3LYP-D3[58] and M06-2X and ωB97X-D.[59] Problems with descriptions of LD interactions in biologically relevant conformers (including sugars) by such class of DFT methods were also identified by Goerigk.[25j] Therefore, it was obvious that only B3LYP optimized geometries of glucopyranosides **1** should be further considered. Our choice was in line with the conclusion of Roy *et al.*[25h] that the density functionals specifically designed to address dispersion behave rather erratically for some systems (but with a tendency to overestimate the strength of LD effects), while B3LYP can describe these interactions as well or better than some specialized functionals.

As to small IHV frequencies found analytically for eight B3LYP-GD3BJ geometries of **1α** and **1β**, we decided to check whether the use of a standard IEF-PCM model of solvation was the most probable cause of such findings (as was suggested by one reviewer). Indeed, the IHV modes resulted from too short vdW radii of lithium atom were found in the DFT study on some Li-containing species.[60] Accordingly, all eight 'wrong' B3LYP-GD3BJ structures were recalculated using the three other vdW atomic radii, namely, UA0 and Bondi (both available in Gaussian 09) and IDSCRF.[61] These latter, isodensity-based SCRF radii, were recently evaluated[61] and applied by the Fang group in mechanistic considerations,[59,62] as a correction of the default IEF-PCM approach implemented within Gaussian 09. The new results such obtained are collected in Table S6†, together with those concerning the precursor UFF radii-based structures. Inspection of this table revealed that the gradual change from UFF to IDSCRF *via* UA0 and Bondi radii gave good results in the majority of cases. Indeed, four positive or two a little negative ω_e values were computed using the IDSCRF radii but a lack of improvement is also found (2 forms). Especially erratic results were obtained for the structure **13787** including an outstanding ω_e value of 12.5*i* cm^{-1} found by use of the radii of Bondi. It should be noted that a high-quality integration grid and convergence threshold were applied in all calculations;[60,63] see Computational. In conclusion, *our results strongly suggest an imperfection of the IEF-PCM/B3LYP-GD3BJ approach*. Indeed, all of these 'wrong' geometries are undoubtedly genuine energy minima because are very similar to their ωB97X-D counterparts (or B3LYP-GD3BJ structures obtained with another vdW radii) showing real vibrational frequencies.[64] Moreover, only the use of the B3LYP-GD3BJ functional provides such wrong results for various radii. Hence, all of the above-discussed IHV frequencies, being well within the range of accepted computational accuracy (~±20[65] or even ~±50[60] cm^{-1}) arising from errors of the numerical integration procedures used in DFT calculations,[63c] are considered to be artificial. Our findings indicate, on the other hand, that further improvements of the existing implicit solvation models are possible.

To circumvent the whole problem concerning the title compounds **1**, a non-classical 'method of gradual exclusion' had to be used to make the analysis tractable. Thus, it was realized that (*i*) an unusual gt_{90} rotamer, which was originally only found for two forms of **1α** in our extensive MM search, can be safely discarded as a critical determinant of related δ_H data. Indeed, the δ_Hs predicted for two anisochronous methylene protons at C6' in the $CHCH_2OAc$ molecular unit, adopting such 'bent' gt_{90} forms, strongly deviate from the observed values by *ca.* –0.7 and +0.7 ppm for the prochiral H6'*R* and H6'*S* protons, respectively. In turn, two vicinal time-averaged *J* couplings within these units, measured for glucopyranosides **1α** and **1β** as $^3J_{H5',H6'S}$ = 2.5 ± 0.2 Hz and $^3J_{H5',H6'R}$ = 4.7 Hz,[66] indicated, in view of the above assumption and our predicted J_{HH} data given below in parentheses, that (*ii*) a contribution of *gg* forms ($^3J_{H5',H6'S}$ ≈ 2.3 Hz, $^3J_{H5',H6'R}$ ≈ 2.2 Hz) to equilibrated mixtures must be approximately twice greater than that of related *gt* forms ($^3J_{H5',H6'S}$ ≈ 1.4 Hz, $^3J_{H5',H6'R}$ ≈ 9.1 Hz), because measured $^3J_{HH}$s are mainly due to motional averaging of such rotamers in solution. This finding was qualitatively consistent with the *gg*/*gt*/*tg* ratio of 53:38:9 and 49:47:4 proposed, respectively, for **2α** and **2β** having an identical sugar part, on the basis of $^3J_{C5,H6}$s measured in C_6D_6

solution.[32] Moreover, (*iii*) a participation of the puckers HC^+ and HC^- of a flexible DHP ring is most likely comparable, as very similar values of δ_H and δ_C were found for the 2a/2b *gem*-dimethyl groups. An analogous conclusion can also be drawn from the X-ray analysis of **1α** showing coexistence of two different half-chair forms in the crystal state.[18] It should be also noted, that all above guidelines (*i*)–(*iii*) were fully in line with considerations of the effect of magnetic anisotropy of the 6'-*O*-acetyl carbonyl group[43a] and an aromatic core of the aglycone (diamagnetic ring current), respectively.

As a result, only eight IEF-PCM(UFF,$CHCl_3$)/B3LYP//6-31+G(*d*,*p*)-optimized structures of every anomer of **1** denoted as forms **1αA** to **1βH** were further studied; their geometries and atomic Cartesian coordinates are listed in Tables S4, S5, S15 and S16†. At this stage, the Grimme's D3 scheme[25c] was *a posteriori* applied to account for impact of weak intramolecular LD effects on related energetic data. More precisely, the total standard Gibbs free energy G_{tot}^{o} of every single form was approximated by a dispersion-corrected G_{DFT-D3}^{o} value considered as including a harmonic DFT contribution, G_{DFT}^{o}, plus a (negative) pairwise interatomic LD correction term E_{disp},

$$G_{tot}^{o} \cong G_{DFT-D3}^{o} = G_{DFT}^{o} + E_{disp},$$

where E_{disp} is the Grimme's semi-empirical B3LYP(G) specific DFT-D3 correction. Such an approach was successfully used in our previous works.[1c,23] Pertinent corrected G_{DFT}^{o}s (= G_{B3LYP}^{o}s), atomic pairwise vdW dispersion terms (DFT-D V3 data),[67] corrected G_{DFT-D3}^{o} data and contributions $p2_i$ (where i = A, B, C…H) of the forms **1αA-1βH** to their equilibrium mixtures in simulated $CHCl_3$ solutions are collected in Tables S7 and S9† together with the $p1_i$ values calculated, according to the Boltzmann distribution law, from the uncorrected G_{B3LYP}^{o}s ('Boltzmann 1' data). For completeness, the initial code names of all 16 finally selected conformers of **1** are also included. Because of the inherent limited accuracy of conventional DFT approaches, the differences in energies (E_{tot}s or G^{o}s) less than the 'chemical accuracy' of 4 kJ mol^{-1} mean comparable thermochemical stability of the predicted structures.[38c,49b] This opinion is consistent with our findings on relative stability of both anomers of **1**. Indeed, the difference in values of $G_{DFT\ 390}$ and $G_{DFT-D3\ 390}$ estimated 'in $CHCl_3$' for their lowest-energy forms **1αA** and **1βB**, amounts to 9.11 and 2.41 kJ mol^{-1}, respectively, whereas ΔG_{390} = 1.87 kJ mol^{-1} follows from an experimental α/β ratio of 36:64 (*vide supra*). These results strongly indicate the need of the usage of dispersion corrections and suggest that the discrepancy in our G_{DFT-D3}s is only 0.5 kJ mol^{-1}. Consequently, energetic orderings gathered in Tables S7 and S9† that resulted from similar thermodynamic data were considered relatively good indicators.

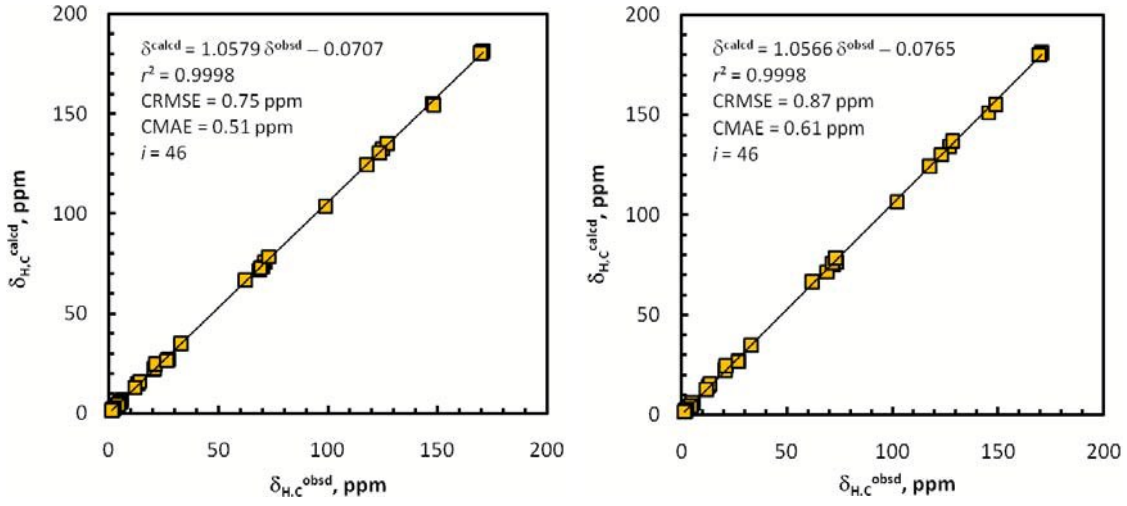


Fig. 5. Scatter plots of DFT computed *vs.* experimental (1:1) $\delta_{H,C}$ data sets for the overall multi-component solution conformation of **1α** (left side) and **1β** (right side); for an additional information see text, Table 1 and Computational details.

In the final stage of this research, the GIAO/DFT-based values of $\delta_{H,C}$s and a few *J* couplings predicted for the individual forms **A-H** of **1α** and **1β** were confronted with respective parameters of NMR spectra measured in solution, by using a linear regression analysis (Computational). Relative populations $p3_i$ of these conformers, roughly known from the foregoing discussion rooted in an NMR experiment were applied as our supplementary and complementary guidelines. The analysis of all structural information indicated that simultaneous fitting of chemical-shift values and some diagnostic $^{n}J_{HH}$ data regarding, respectively, the *gem*-dimethyl and C*H*C*H*$_2$OAc units in both glucosides **1**, was of crucial importance. The findings from such combined experimental-theoretical approach supported by the statistical treatment are shown graphically in Fig. 5 and summarized in Table 1. All three relevant statistical indicators (r^2, CRMSE, and CMAE; see Computational for details) are given in the plots as estimates of the reliability of results.

Table 1. Relative abundances of the forms **A-H** of glucosides **1α** and **1β** according to three different 'theory *vs.* experiment' considerations of the energetic (ΔG) and NMR ($\delta_{H,C}$/J_{HH}) data.[a]

i	**A**	**B**	**C**	**D**	**E**	**F**	**G**	**H**
α-anomer (**1α**)								
C5–C6 rotamer	*gg*	*gt*	*gt*	*gg*	*gg*	*gt*	*gt*	*gg*
HC pair [b]	I–	II+	III–	IV–	I+	II–	III+	IV+
$p1_i \times 100$, Boltzmann 1, % [c,d]	18.8	15.9	14.8_5	13.8	13.7	9.1	7.9	5.9_5
$p2_i \times 100$, Boltzmann 2, % [c,e]	*24.6_5*	*11.3*	*11.1*	*17.2*	*17.7_5*	*4.7_5*	6.4	6.8
$p3_i \times 100$, Boltzmann 3, % [c,f]	21.9	4.3	4.6	22.2	17.2	4.8_5	3.0	21.8_5
$p4_i \times 100$, Boltzmann 4, % [c,g]	21.4	4.9	4.5	21.3	19.8_5	5.0	3.3	19.9
$p5_i \times 100$, DFT/NMR data, %	*19*	*10*	*4*	*15*	*19*	*8*	10	15
$p6_i \times 100$, MP2/NMR data, % [h]	19.5	10	5	15	18	7	10	15.5
β-anomer (**1β**)								
C5–C6 rotamer	*gt*	*gg*	*gg*	*gg*	*gt*	*gt*	*gt*	*gg*
HC pair [b]	I–	II–	II+	III+	I+	IV–	IV+	III–
$p1_i \times 100$, Boltzmann 1, % [c,d]	26.5	21.0	12.8	10.4_5	9.7	7.5	6.5	5.5
$p2_i \times 100$, Boltzmann 2, % [c,e]	*19.2*	*27.0*	*15.6*	*13.6_5*	**7.6**	5.6	4.6	6.7
$p3_i \times 100$, Boltzmann 3, % [c,f]	4.3	18.3	15.6	22.3	4.6	5.2	5.6	24.1
$p4_i \times 100$, Boltzmann 4, % [c,g]	4.8	20.1_5	16.7	21.7	5.4	5.6	5.6	20.1
$p5_i \times 100$, DFT/NMR data, %	*14*	*14.5*	*10.5*	*20*	**16**	0	5	20
$p6_i \times 100$, MP2/NMR data, % [h]	14	14.5	10.5	20	**16**	0	5	20

[a/] The greatest divergence in the p_i populations is given in bold type. [b/] Corresponding HC pairs of DHP rings (with the **+/−** sign of θ) for the same *gg* or *gt* form. [c/] For all details, see Tables S7-S10†. [d/] Without the dispersion correction. [e/] With the dispersion correction. [f/] Without the correction for $ZPVE_{DFT}$. [g/] With the correction for $ZPVE_{DFT}$. [h/] For cut-off subsets of the $\sigma_{H,C}$ data (see text).

Inspection of Table 1 (and Tables S11 and S12†, with the *p*1 and *p*2-based values of selected NMR data, respectively) reveals that the use of dispersion corrected G^os really led to a much better agreement between populations of single species accessed from the energy *vs.* DFT/NMR data, at least for the forms **1αA-1αF** and **1βA-1βD**; see italicized figures relating to *p*2 and *p*5 data. This result for the studied seven/eight-conformer objects is in full accord with similar conclusion drawn from our previous study limited to the three-component systems.[1c] The 3D shapes of the most privileged *gg* forms, *i.e.,* **1αA** and **1αE** ($p3 = 0.19$) as well as **1βD** and **1βH** ($p3 = 0.20$), are depicted in Fig. 6. This finding seems to indicate that the aglycone part of both anomers adopts mainly the same orientation with respect to their glycone moieties (the Me group at C5 close to O5'). All conformers of **1** with percentage populations are shown in Figs. S10 and S11†. It should be noted that the magnitude of r^2 was not decisive in the analysis, because only very small changes in the magnitude of this correlation indicator were found for **1αA-1βH** on going from the *p*1 (or *p*2) to *p*5 data (Tables 2, S11, and S12†). In sharp contrast, a great change (from ~1.0 to ~1.8) in the *gg*/*gt* rotamer ratio was observed on

coming from the *p*1 to *p*2 results, strongly suggesting that structures with the $CHCH_2OAc$ unit in the *gg* conformation are favored by LD forces.

Overall, only slightly weaker correlation between the predicted and experimental $\delta_{H,C}$ sets was found for **1β**. Indeed, the greatest discrepancy in *p*2s/*p*5s (~8%) was obtained for **1βE** (Table 1). Nevertheless, only for these NMR-based populations very small differences in the simultaneously analyzed data of δ_H, δ_C and $^nJ_{HH}$ were found in a laborious but critical step in achieving very good reproduction of the observed values of chemical shifts of 2a/2b *gem*-dimethyl groups and J_{HH}s in the C*H*CH_2OAc unit. Additionally, population-averaged values of the other computed *J* data discussed in the text, *i.e.,* $^2J_{H3,H4}$ = 6.60, $^4J_{H1',H3'}$ = –0.45, $^4J_{H1',H5'}$ = –0.72, $^4J_{H2a,H2b}$ = 0.46 and $^1J_{C1',H1'}$ = 168.79 Hz for **1α** as well as $^2J_{H3,H4}$ = 6.59, $^4J_{H2a,H2b}$ = 0.47 and $^1J_{C1',H1'}$ = 161.87 Hz for **1β**, were obtained in good agreement with those found experimentally; for a scatter plot of selected relationships $J_{HH}^{calcd\ av}$ *vs.* J_{HH}^{obsd} (r^2 = 0.99900) see Fig. S14†.

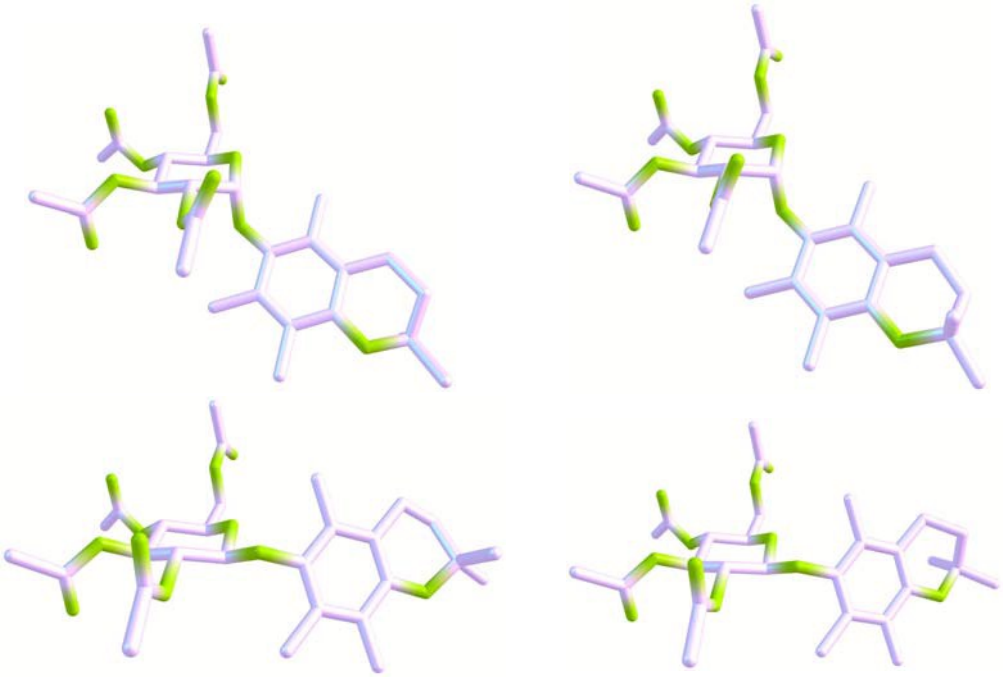


Fig. 6. Chemcraft 3D drawings of the four *gg* forms of glucopyranoside **1** favored in solution according to the DFT/NMR results: **1αA** (19%, top, left), **1αE** (19%, top, right), **1βD** (20%, bottom, left) and **1βH** (20%, bottom, right). Two different ring-flipped forms of a DHP moiety are visible for every anomer. For clarity, all hydrogen atoms have been omitted.

Moreover, the structure **1βA**, observed as its enantiomeric form in the crystal of **1β**,[42] was relatively strongly privileged in $CHCl_3$ solution (*p*5 = 0.14), in sharp contrast to the case with the α-anomer. Indeed, close inspection of the crystal structure of **1α**[18] suggests that the coexistence of the four species with a 'bent' gt_{90} rotamer of the $CHCH_2OAc$ segment (different from those located in our MM search) in the unit cell is due to crystal packing effects largely dominated by intermolecular contacts of the CH···HC type,[55] involving *inter alia* the Me group of this unit interacting with 2a/2b-*gem*-dimethyl groups of a neighboring molecule (see also above). A great similarity between angles φ, ψ and χ in both these main gt_{90} conformers and their related non-physical solution M06-2X structures optimized with the allowance for LD forces supports this conception (Table S4†, the forms **14229** and **14913**). As a result, 'extended' *gt* and, especially, *gg* rotamers of this molecular unit in both glucosides **1** under study are favored in the solution state.

On the other hand, a definitive and unambiguous assignment of the $^1H/^{13}C$ NMR signals of 2a/2b *gem*-dimethyl groups was simultaneously acquired in our analysis. Pertinent chemical-shift values are only a little more differentiated for **1α**, but a *downward* Me substituent labeled 2a was always found to

resonate downfield of its *upward* counterpart 2b (Fig. 1, Table 2); the spatial relationship between these groups adopted throughout this paper is the same as used before.[13a] Therefore, one cannot tell about methyl groups C2a and C2b as being located, respectively, in *equatorial* and *axial* positions as was considered previously.[13a] In this case each of these two Me groups occupy both such orientations during low-energy interconversions (rapid in the NMR time-scale) between two different half-chair forms of a DHP ring.[13a,68] Interestingly, an experimental $\Delta\delta_C/\Delta\delta_H$ ratio amounts to 18.5 ± 0.2 for both anomers and also the average values of δ are very similar, $\delta_H = 1.285 \pm 0.002$ ppm and $\delta_C = 26.78_5 \pm 0.02_5$ ppm. All above facts indicate that the structural features and dynamics of the DHP part of both anomeric glucosides **1** in solution are comparable.

It should also be noted that the *gg*/*gt* rotamer ratio determined for **1α** is a little greater than that found for **1β** [the α/β ratio (of both *gg*/*gt* ratios) ~1.1_5], see Table 2, similarly as was estimated[32] for the pair **2α** and **2β** having the same sugar moiety (α/β ~1.3). In light of these results, more recent literature data[43c] suggesting the *gg*/*gt* ratio of 0.61 and 0.52 for **2α** and **2β**, respectively, are questionable, but the associated α/β ratio ~1.2 is correct.

An inherent uncertainty of the finest GIAO/DFT-based $p5_i$ values is difficult to estimate, due to possible summation and/or cancellation of errors in two subsequent computations of geometries and chemical shifts (or ΔG_{DFT-D3}°s). The differences between the $p2_i$ and $p5_i$ results found for **1** (Table 1) suggest that such uncertainty is in the order of 4-7%, under the assumption of perfect correctness of $p2_i$ data. But, one should remember about a modest accuracy of typical Δ*G*s and so about relatively large errors in calculations of $p1_i$s and perhaps also, to some extent, $p2_i$s. Thus, it seems that the uncertainty in question is comparable with that reported previously for the best example of three-component systems studied analogously (most likely <5%).[1c] So it was concluded that values of $p2_i$s and $p5_i$s are consistent with each other within their errors; however, very good agreement with the NMR experimental observations was found for the latter dataset only. Hence, one can invoke again the concept of superiority of the '*solution match criterion*' over '*thermodynamic criterion*', stressing simultaneously that accounting for weak LD forces in calculations of Δ*G*s and thus Boltzmann distributions is mandatory in all such cases. A very similar conclusion was drawn earlier.[1c]

Table 2. Selected $^1H/^{13}C$ chemical shifts [ppm] and J_{HH} couplings [Hz] relating to the 2a/2b *gem*-dimethyl and C*H*CH_2OAc units of forms **A**-**H**, respectively, found for the GIAO NMR based *p*3 (δ_Ks and $J_{H,H}$s) and *p*6 (only δ_Ks) data.

	α-anomer (**1α**)				β-anomer (**1β**)			
Nucleus *j*	**Exp.**	**Calcd** [a]	**Scaled** [b]	**$-(\delta^{scaled} - \delta^{obsd})$**	**Exp.**	**Calcd** [a]	**Scaled** [b]	**$-(\delta^{scaled} - \delta^{obsd})$**
H2a	1.298	1.231 [c]	1.231	0.067	1.293	1.227 [c]	1.303	–0.010
H2b	1.268	1.199	1.201	0.067	1.281	1.215	1.292	–0.011
C2a	27.04	26.85	25.45	1.59	26.92	26.69	25.33	1.59
C2b	26.48	26.25	24.88	1.60	26.70	26.46	25.12	1.58
H2a	1.298	1.342 [d]	1.056	0.242	1.293	1.343 [d]	1.004	0.289
H2b	1.268	1.312	1.029	0.239	1.281	1.332	0.994	0.287
C2a	27.04	29.34	27.02	0.02	26.92	29.20	26.93	–0.01
C2b	26.48	28.77	26.49	–0.01	26.70	28.95	26.69	0.01
Coupling H,H	**Exp.**	**Calcd** [e]	**–**	**$-(\delta^{scaled} - \delta^{obsd})$**	**Exp.**	**Calcd** [e]	**–**	**$-(\delta^{scaled} - \delta^{obsd})$**
$^3J_{H5',H6'S}$	2.34	2.06	–	0.28	2.74	1.98	–	0.76
$^3J_{H5',H6'S}$	4.68	4.46	–	0.22	4.72	4.59	–	0.13
$^2J_{H6'R,H6'S}$	(–)12.37	–12.63	–	0.26	(–)12.18	–12.51	–	0.33
DFT/NMR [f]	$r_{C/H}^2 = 0.99983$ *gg*/*gt* = 2.13 CRMSE = 0.75 CMAE = 0.51 $\delta^{calcd} = 1.0579\ \delta^{obsd} - 0.0707$				$r_{C/H}^2 = 0.99977$ *gg*/*gt* = 1.86 CRMSE = 0.87 CMAE = 0.61 $\delta^{calcd} = 1.0566\ \delta^{obsd} - 0.0765$			
MP2/NMR [f]	$r_{C/H}^2 = 0.99930$ *gg*/*gt* = 2.13 CRMSE = 0.73 CMAE = 0.58 $\delta^{calcd} = 1.0784\ \delta^{obsd} + 0.2031$				$r_{C/H}^2 = 0.99922$ *gg*/*gt* = 1.86 CRMSE = 0.79 CMAE = 0.67 $\delta^{calcd} = 1.0746\ \delta^{obsd} + 0.2641$			

[a/] $\delta_{K,j}^{calcd} = \sigma_{K,TMS} - (p3_A \cdot \sigma_{K,A,j} + p3_B \cdot \sigma_{K,B,j} + \ldots + p3_H \cdot \sigma_{K,H,j})$, K = H or C. [b/] $\delta_{K,j}^{scaled} = (\delta_{K,j}^{calcd} - b)/a$, for the least squares linear fitting values of the slope *a* and intercept *b*, see below and Fig. 5. [c/] GIAO/DFT data-based results. [d/] GIAO/MP2 data-based results. [e/] $J_{HH} = p3_A \cdot J_{A,HH} + p3_B \cdot J_{B,HH} + \ldots + p3_H \cdot J_{H,HH}$ (found at both theory levels for the DFT-level *J*-data); see Computational. [f/] A binuclear ($\delta_{H,C}$ 1:1) regression analysis was applied, see also Computational.

The foregoing results based on the DFT data were finally compared with those arising from the total electronic-nuclear energies, E_{tot}s, and GIAO predictions of σ_Ks obtained for solutions of **1αA-1βH** at the MP2[69]/6-311+G(*2d*,*p*) and MP2/6-311G(*d*,*p*) level, respectively. As to energy data and thus conformer populations *p*3 and *p*4 found from relative energies ΔE_{tot} and ΔE_0 (Tables 1, S8 and S10†),[70] the new results are generally consistent with DFT findings, with the exception of cases of **1αB**, **1αH**, **1βA**, and **1βH** (Table 1, numbers underlined). But also in these instances, the MP2 data-derived results are in agreement with the trends observed on going from *p*1s to *p*2s within the limits of inherent errors of both theoretical models.[71,72] Also 'mean' populations found for related *p*2/*p*4 pairs, namely 8.1, 13.3_5, 12.0 and 13.4%, are in line with the *p*5 DFT/NMR data. As previously stated, the greatest discrepancy between *p*4 and *p*6 values is found for **1βE**. Therefore, one can then consider based on such new data (Table 1, *p*3s and *p*4s) that the results of MP2 calculations are qualitatively consistent with the DFT-D3-type intramolecular LD attractions in the systems **1α** and **1β** estimated here. Moreover, according to the aforementioned data, it seems that the inclusion of DFT-level ZPVE terms in calculations of ΔE_0s and so *p*6 values was fully justified, despite some warnings on this topic concerning systems with the relatively flat potential energy hipersurfaces.[73]

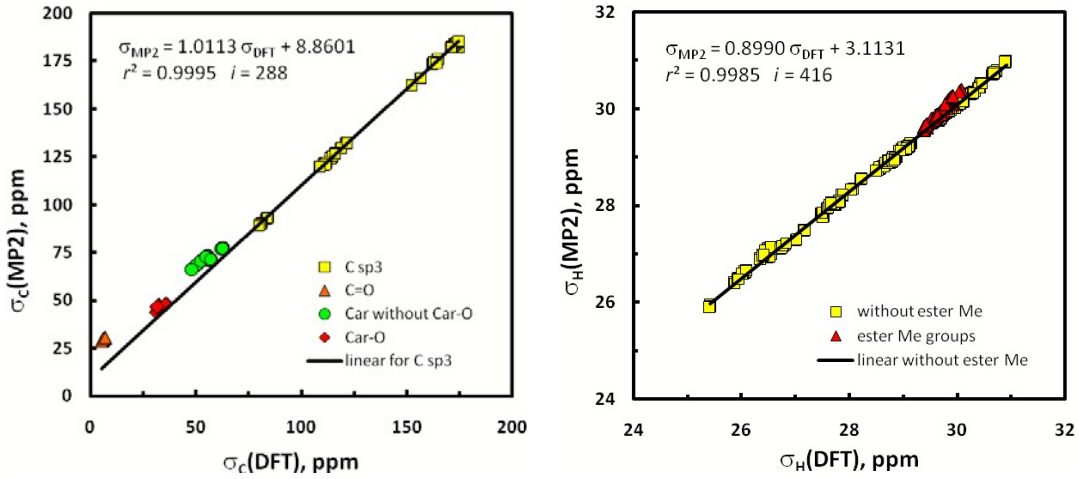


Fig. 7. Regression plots of the relationships between the IEF-PCM(UFF,$CHCl_3$)/GIAO/MP2 and DFT-mPW1PW91 calculated isotropic shieldings concerning the same geometries of the forms **1αA-1βH**: (**a**, left side) ^{13}C nuclei and (**b**, right side) ^{1}H nuclei; for an additional information see text and Computational details.

The situation with the results of GIAO/MP2 calculations[74] is more complicated. Generally, these data seem to be by far less exact than related DFT-level findings concerning the same geometries and being in good accord with the empirical observations. Against very good correlation between σ_Cs computed at the MP2 *vs.* DFT level, awaited in light of the pioneering results of Wiberg (r^2_C = 0.9994),[75] only a good relationship was found for all ^{13}C nuclei in the 16 conformers under study (r^2_C = 0.9977 for $i = 16 \times 28$ unique nuclei);[76] the correlation between all σ_Hs is still weaker (r^2_H = 0.9946, $i = 16 \times 38$). Evidently, both models of chemistry provided different GIAO predictions for ^{13}C nuclei involved in π-systems (especially in the ester C=O bonds). Indeed, exclusion of all sp^2 hybridized carbons in the σ_C set give r^2_C = 0.99953 (Fig. 7a); four different clusters of data points due to sp^2 carbons are worth mentioning. The same is also true to some degree with σ_Hs, and omission of all protons of the methyl ester groups experiencing an anisotropic effect of neighboring C=O groups lead to r^2_H = 0.99845 (Fig. 7b). Therefore, only use of two such cut-off subsets of the σ values in subsequent binuclear $\delta_{H,C}^{calcd}$(MP2) *vs.* $\delta_{H,C}^{obsd}$ correlations important for this investigation was fully legitimate (for related plots, see Fig. S15†). But the MP2/NMR populations so obtained ($p6_i$ values, Table 1) are a little less reliable owing to a lack of some data points – as previously stated, the best

reproduction of 'diagnostic' patterns of δ_Ks concerning the *gem*-dimethyl groups at C2 and $^nJ_{HH}$s around C6' was of crucial importance. Slightly changed $p6_i$s were such obtained for **1α**, but all attempts to correct the $p5_i$ data used as tentative starting values for **1β** were unsuccessful.

On the whole, a satisfying agreement with the earlier DFT/NMR results was found (Table 1). The discrepancies between the DFT and MP2-derived values of $\Delta\delta$ (= $\delta^{scaled} - \delta^{obsd}$) concerning the 2a/2b Me groups arise mainly from different slopes of related best-fit lines (Table 2; *cf.* Figs. 5 and S15†). Such $\Delta\delta$ data obtained for *p*3 and *p*4 abundances (Tables 2, S13 and S14†) are less consistent, but those found for the *p*4s are better. Also the *gg*/*gt* ratios improve on going from *p*3 to *p*4 values (**1α** 5.0→4.7, **1β** 4.1→3.7). However, what must be emphasized here is that all these data are incompatible with the NMR spectroscopic observation (*gg*/*gt* ~2, *vide supra*). Interestingly, the reverse trend in *gg*/*gt* is observed on passing from *p*1 to *p*2 data evaluated from the DFT results (**1α** 1.1→2.0, **1β** 1.0→1.7; Tables S11 and S12†). *Thus, is it possible that dispersive attractions (?)[55] between H5' and the two H6' atoms in gg rotamers of the CHCH$_2$OAc fragment of systems* ***1*** *(see Fig. 3) are favored too much in MP2 and underestimated in B3LYP treatments?* In summary, one can consider that the results emerging from MP2 calculations confirm the earlier DFT results, though certain disagreements between them (and with the experiment) were also recognized. Particularly interesting are the foregoing discrepancies in σ_Cs predicted at both levels of theory.

For some other findings, it was recognized that the large differences $\Delta_{H5} = \delta_{H5\alpha} - \delta_{H5\beta}$, observed for anomeric pairs of several *O*-glycosides of PMC (*vide supra,* see also Table S2†), must arise from an aromatic ring-current effect of the constituent chroman system. Indeed, inspection of molecular representations of all forms **A-H** of **1α** (Figs. S10†) revealed that their hydrogen atoms at C5' are situated within the deshielding cone produced by circulating π-electrons. By contrast, a relative small shielding of both H6' protons (in relation to those occurred in **1α**) is suggested on the basis of 3D drawings of all forms of **1β** (Figs. S11†), in full agreement with the NMR data.

Conclusions

In this combined theoretical/experimental study two highly flexible glucoside derivatives of PMC (a model compound of α-tocopherol) were used to test several current calculational protocols accessible for predicting the overall shapes of multi-conformer systems and population-weighted averages of their NMR parameters based on high-quality spectroscopic data. A special emphasis was given to accounting for the impact of intramolecular LD effects on the geometries and relative Gibbs free energies (ΔGs) of various forms coexisting in solution. Detection of a few small $^4J_{HH}$ coupling constants in both molecules is also worth mentioning.

Of many possible single conformers of glucopyranosides **1α** and **1β** localized in initial Monte Carlo MM searches, only twelve amongst them were finally recognized in quantum-chemical calculations to contribute significantly (≥10%) to related conformational mixtures in solution, where solvent effects on geometries and NMR spectral properties of the analyzed solutes were mainly simulated with an IEF-PCM(UFF,$CHCl_3$) approach of implicit solvation. Simultaneous matching of computed *vs.* observed NMR chemical-shift sets applying the binuclear ($\delta_{H,C}$ 1:1) linear regression analysis was considered the best procedure in disentangling the conformational preferences of these systems. The presence of their 2a/2b-*gem*-dimethyl and $CHCH_2OAc$ structural units, as sensitive intrinsic magnetic probes for detecting time-averaged spatial arrangements of the atom arrays in their nearest electronic environments (local solute geometries), was recognized as being of crucial importance for achieving good reproduction of solution NMR spectra of both anomers.

Regarding the molecular structure of **1α** and **1β**, the four DFT functionals including three with *a priori* corrections for attractive LD forces (M06-2X, ωB97X-D and B3LYP-GD3BJ) gave different geometries. The best results were found with B3LYP, while the two last specialized DFT methods afforded *tg* rotamers of the $CHCH_2OAc$ fragment instead of related *gg* forms in contradiction with the experiment. The advantage of an application of IDSCRF over default UFF radii in the IEF-PCM simulations of solvation was simultaneously shown for several B3LYP-GD3BJ optimized structures having one small imaginary vibrational frequency. All these findings strongly suggest that functional ωB97X-D and especially B3LYP-GD3BJ are rather not suitable for modeling the ground-state geometry of highly flexible molecules. Moreover, some serious problem with the IEF-PCM/B3LYP-GD3BJ approach was found.

The Gibbs free energies of individual forms of **1α** and **1β** optimized by the B3LYP/6-31+G(*d*,*p*) method were subjected to vdW (DFT-D3) corrections for LD effects to give respective G_{DFT-D3}s. These latter values gave (*via* Boltzmann statistics) estimated populations of single forms in the solution mixtures ($p2_i$s) being in much better agreement with NMR data-based populations ($p5_i$s) than those calculated for uncorrected G_{DFT}s ($p1_i$s). Related $p2$ and $p5$ values were found practically equivalent within their error limits, but only these latter values gave very good agreement with the observation. Very similar conformer populations were also derived from the MP2/NMR data ($p6_i$s). These findings confirm the need for *post factum* running of LD corrections in DFT studies of this kind.

A *gg*/*gt* rotamer ratio of ~2.0 was established for the $CHCH_2OAc$ fragment of both glucosides on the basis of DFT data (the MP2 energetic results give a considerably overestimated value of ~4.2). Also such ratio, estimated from the G_{DFT-D3} data, was much better than that found from initial G_{DFT}s (~1.8 *vs.* ~1.0). The more compact gt_{90} rotamer of this unit was not recognized in solution and so its presence in the crystal structure of **1α** originates evidently from packing effects. In contrast, its *gt* rotamer identified in the crystal of **1β** was found as one of the five predominant forms in solution. It was also established that the differences $\Delta_{H5} = \delta_{H5\alpha} - \delta_{H5\beta}$, observed for anomeric pairs of some *O*-glycopyranosides of PMC, are due to the aromatic ring-current effect of a chroman skeleton. Hence, this parameter is proposed as a determinant of stereochemistry at anomeric centers in molecules of this kind. All the main findings of this work were confirmed by additional calculations done at the MP2 level. Simultaneously, some interesting discrepancies in values of σ_Cs predicted at both theory levels were recognized. One can suppose that with the applied (or equivalent)[77] MM/DFT methodology and careful analysis of results, it is possible to find all, or at least the huge majority, of the low-energy conformers of various other small- to medium-sized flexible molecules. Hence, we believe that our results prove useful for guiding similar joint NMR spectroscopic/DFT computational studies on further multi-conformer systems in solution, especially those having the sugar moiety as a structural motif.

Computational details

Geometry optimization, vibrational frequency and energy calculations

A stochastic conformational search for minima on the potential energy hypersurfaces of the objects **1α** and **β** was performed with the Global-MMX (GMMX) subroutine built into the PCMODEL 8.5 package.[44,78] Specifically, a mixed MM protocol,[79] based on Monte-Carlo (MC) procedures used originally in the BAKMDL program,[80] was employed in which randomly selected atoms of the semi- and saturated (hetero)cyclic rings and all of the seven rotatable bonds were randomly moved in the Cartesian[70,81] and dihedral angle[82] space and energies of such formed species were subsequently minimized within the MMX (1986) force field.[83] About 40 cycles of GMMX calculations, each

embracing 5000 of MC searching steps, were performed for every molecule with the bulk relative permittivity (dielectric constant, ε) of 4.71[34] used for a rough simulation of the $CHCl_3$ environment. A search was continued until ~180 unique energetically lowest energy lying structures of each anomer were generated within an arbitrary chosen 25.1 kJ mol^{-1} energy window. The so-obtained MMX models were then subjected to a gradient geometry optimization, initially at HF/3-21G[84] and next (after sorting and removing of duplicates) at HF/6-31G(*d*) levels, by using the Gaussian 09 suite of electronic structure programs.[34] All types of geometric motifs of various occurring rotameric forms were recognized in this way. Initial MMX structures of the seven not originally located conformers were built without any changes in atom numberings through adequate modifications of the partially (or fully) optimized geometries of the relevant closely related forms,[47] by using Hyperchem[46] (MM+ force field);[45] for all details see footnotes to Tables S4 and S5†. The MM+ calculations were followed by MMX optimizations in these additional cases. It should be noted that very large differences in energetic ordering of the input MM models of **1α** and **β** (established *via* their MMX energies) and pertinent HF/3-21G optimized structures (*via* the ΔE_{tot} data) were generally found; similar situation was observed previously.[23b]

Final geometry refinement of the 'best' structures was carried out at the double-ζ (DZ) valence quality level of theory using the hybrid B3LYP[36a-c] exchange-correlation functional, as implemented in the Gaussian code,[2c,36d] in conjunction with the 6-31+G(*d,p*) basis set recommended for DFT calculations of energy data,[36f] especially for the systems with lone electron pairs on heteroatoms.[85] For the sake of accuracy, the 'Tight' SCF and Opt convergence criteria were used in all computations.[1c,60,63] Moreover, a fine-pruned (150,590)[86] numerical integration grid having 150 radial shells and 590 angular points per shell was always selected applying the Int(Grid=150590) keyword,[63d-f] because of soft modes coming from dynamic phenomena of methyl groups rotations.[34] Simultaneously, an attempt to evaluate solvent influences on the solute structures and properties was made within an equilibrium solvation protocol[20b] of an integral equation formalism-polarizable continuum model (IEF-PCM)[19,20] of solvation, by employing the UFF atomic radii when constructing the solvent cavity and other default parameters. Analogous optimizations in implicit $CHCl_3$ solvent were also carried out with the use of a 6-31+G(*d,p*) basis set and three specialized DFT functionals *a priori* corrected for contributions of LD effects, namely, B3LYP-GD3BJ (*i.e.,* B3LYP with addition of the D3 version of Grimme's dispersion[25c] with Becke-Johnson damping[25e] [Gaussian keyword: B3LYP/base EmpiricalDispersion=GD3BJ],[34] M06-2X,[52] and ωB97X-D.[53] Some additional structures were also optimized with IEF-PCM/B3LYP-GD3BJ applying the three other atomic radii, *i.e.,* UA0 and Bondi (both available in Gaussian 09) and IDSCRF[61] (see also text). Fully-relaxed geometries of 16 finally considered forms of **1** found at the IEF-PCM(UFF,$CHCl_3$)/B3LYP/6-31+G(*d,p*) level are given in Tables S15 and S16†, while their 3D shapes are depicted in Figs. S10 and S11†, using graphical representations created with the ChemCraft program.[87]

Moreover, vibrational wavenumbers ω_e were always evaluated in the rigid rotor-harmonic oscillator-ideal gas approximation of vibrational modes that was used in the frame of the same DFT method, to verify whether the located stationary points represented true minima (N_{imag} = 0) on the Born-Oppenheimer ground-state energy hypersurfaces of analyzed structures and to determine their unscaled ZPVE corrections and Gibbs free energies, G^o_{298}s, at standard ambient temperature and pressure (298.15 K, p = 1 atm), *i.e.,* close to the NMR recording temperature of 302 ± 2 K. Finally, all of these G^o_{298} data were corrected for vdW dispersion effects (LD forces)[1c,24,25] as was explained in the text, by using respective B3LYP(G) specific D3 Grimme's DFT-D V3 correcting terms[25c] computed with the ORCA package.[67]

In addition, individual total energies, E_{tot}s, of all forms **1αA-1βH** were single-point calculated[74] by the second-order Møller-Plesset (MP2) perturbation method[69] with the 6-311+G(*2d,p*) basis set of

triple-ζ (TZ) valence quality.[70] These computations were additional jobs in the MP2 runs carried out as is described below in the section on NMR spectra predictions.

For assessing relative abundances of individual forms in the conformational equilibria in solution, the fractional Boltzmann population (mole fraction, p_i) of each entity was found using the Boltzmann distribution function, $p_i = e^{-\Delta Gi°/RT}/\Sigma_j e^{-\Delta Gj°/RT}$, where j is the number of species in thermal equilibrium, R is the ideal gas constant, T is the system absolute temperature set to 298.15 K, and $\Delta G_i°$ is the ΔG value of the ith form relative to the energy of the most stable conformer. For the MP2-level results, ΔE_0s were used instead of ΔG values in the calculation of *p*3 and *p*4 data.[70]

Prediction of NMR spectra

Single-point GIAO[2] formalism-based computations of isotropic NMR chemical shielding constants, σ_Ks, were carried out at the IEF-PCM(UFF,$CHCl_3$)/mPW1PW91/6-311+G(2*d*,*p*)[35] level on the IEF-PCM(UFF,$CHCl_3$)/B3LYP/6-31+G(*d*,*p*) computed structures, by using Gaussian 09. Our approach[88] was similar to that used by the Tantillo research group,[3c,35b,c] however, these authors applied another solvent continuum model and used the gas-phase instead of (probably much better)[89] the solution-phase optimized solute structures used here. According to classical tetramethylsilane (TMS) based protocol, the relative chemical shift, δ_K, of a given nucleus K in each molecular entity is defined as δ_K^{calcd} [ppm] = $\sigma_K^{ref} - \sigma_K^{calcd}$. For so predicted ^{1}H and ^{13}C NMR spectra, σ_K^{ref} is equal to 31.7023 and 186.9100 ppm, respectively, as was computed in simulated $CHCl_3$ solution – analogously as mentioned above – for the exact T_d symmetry[90] molecule of TMS as a dual-reference δ_K standard. Several other combinations of functional [B3LYP-GD3BJ,[25c,e,34] M06-2X[52] or ωB97X-D[53] (first step) and mPW1PW91[35a] (second step)] and basis set [6-31+G(*d*,*p*) (first step) or 6-311+G(2*d*,*p*) (second step)] were used in additional GIAO NMR predictions. All of these solution-state calculations were performed on corresponding structures fully pre-optimized at the DZ quality level, see also text. Moreover, supplementary[74] time-consuming GIAO predictions of σ_Ks were performed at the IEF-PCM(UFF,$CHCl_3$)/MP2/6-311G(*d*,*p*)//IEF-PCM(UFF,$CHCl_3$)/B3LYP/6-31+G(*d*,*p*) level for all 16 forms **1αA-1βH** in order to verify the correctness and internal consistency of the GIAO/mPW1PW91 results mentioned above, see text and also Figs. 7 and S15†. These MP2 runs were computationally very demanding tasks. After several initial tests, we were able to do a single GIAO NMR calculation in 8-9 days, by using 24 processors (2.50 GHz), 128 GB of memory, and at least 7.2 TB of scratch disk space for temporary storage of data. The σ_K^{calcd}s obtained in all of these cases were as above referred to TMS applying σ_K^{ref} terms evaluated at the same computational level; σ_K^{ref} (MP2) of 31.8587 and 198.8873 ppm, respectively.

In addition, some indirect couplings, $^nJ_{KL}$, were single-point computed for $CDCl_3$ solutions of **1** at the IEF-PCM(UFF,$CHCl_3$)/B3LYP/IGLO-II level[1c,91] with Gaussian 09. An extended NMR property-oriented IGLO-II basis set of Huzinaga modified next by Kutzelnigg and coworkers (also known as the HII or BII set)[92] and widely used in predicting J_{KL} data,[91,93] was downloaded from the EMSL Basis Set Library.[94] The five so-called pure *d* basis functions were employed for non-hydrogen atoms in all NMR calculations mentioned above.

The GIAO computed σ_Hs of each of the three mutually exchanging hydrogen atoms in the Me groups were arithmetically averaged to produce a single σ_H (or δ_H) value for each Me group as a whole; the same concerns also the two methylene groups of the highly flexible DHP rings. A linear regression analysis of the relationships between the predicted and observed NMR parameters (δ_Ks, in particular) was achieved by a least-squares method; see also footnote *b* to Table 2. More precisely, the calculated data were plotted as usual[1,23b,37] against experimental ones, however binuclear[1a,37,38] 1:1 correlations, $\delta_{H,C}^{calcd}$ *vs.* $\delta_{H,C}^{obsd}$, were applied instead of two separate classical mononuclear

relationships. Such an associated H,C approach was strongly suggested by the analysis of the problems entailed *inter alia* in our previous study dealing with multiple (>2) conformers,[1c] in which application of δ_Cs for assessing populations of the single forms in solution was unsuccessful. The case of the superiority of structural results coming from the GIAO-derived δ_Hs over those from related δ_C data was reported by Koskowich *et al*.[95]

The three relevant statistical metrics, *i.e.,* a square of the Pearson correlation coefficient (r^2), the corrected root-mean-square error [CRMSE equal to RMSE[49b,95,96] with the value^{scaled} data applied instead of the value^{calcd} ones) and the corrected mean absolute error [CMAE,[97] defined as (Σ_i|value^{scaled} – value^{obsd}|)/number of comparisons *(i*)] were used throughout the paper as estimates of uncertainties of the results. The greater value of r^2 (also called coefficient of determination and showing correlation significance) was considered as an indication of better adjustment of correlated data sets. All statistical analysis was performed using an MS Excel 2010 spreadsheet.

Acknowledgements

R.B.N. is grateful to Prof. Laurence A. Nafie (Syracuse University, USA) for the stimulating exchange of letters regarding the VCD spectra. He also thanks to Prof. De-Cai Fang (Beijing Normal University, China) for providing the SCRF-RADII program and to Dr. Piotr Matczak (University of Łódź) for help in installing this software. This work was supported by computer facilities and the Gaussian 09 software in the Academic Computer Centre CYFRONET (AGH – University of Science and Technology, Kraków) through computational grants Nos. MNiSW/SGI3700/UŁódzki/057/2010 and MNiSW/Zeus_lokalnie/UŁódzki/016/2013 (to R.B.N.). R.B.N. is also indebted to the staff of CYFRONET for their assistance in conducting MP2 calculations on the Prometheus supercomputer (within PL-Grid). The authors also thank two anonymous reviewers for their criticism, valuable comments and suggestions of some additional calculations, which results helped to improve the paper.

Notes and references

1 (a) E. Michalik and R. B. Nazarski, *Tetrahedron*, 2004, **60**, 9213–9222; (b) R. B. Nazarski, *J. Phys. Org. Chem.*, 2009, **22**, 834–844; (c) R. B. Nazarski, B. Pasternak and S. Leśniak, *Tetrahedron*, 2011, **67**, 6901–6916.

2 (a) R. Ditchfield, *Mol. Phys.,* 1974, **27**, 789–807 and refs. therein; (b) K. Wolinski, J. F. Hilton and P. Pulay, *J. Am. Chem. Soc.,* 1990, **112**, 8251–8260; (c) J. R. Cheeseman, G. W. Trucks, T. A. Keith and M. J. Frisch, *J. Chem. Phys.*, 1996, **104**, 5497–5509.

3 (a) T. Helgaker, M. Jaszuński and K, Ruud, *Chem. Rev*., 1999, **99**, 293–352; (b) See also some chapters in *Calculation of NMR and EPR Parameters. Theory and Applications*, ed. M. Kaupp, M. Bühl and V. G. Malkin, Wiley-VCH Verlag GmbH & Co. KGaA, Weinheim, 2004; (c) M. W. Lodewyk, M. R. Siebert and D. J. Tantillo, *J. Chem. Rev.*, 2012, **112**, 1839–1862.

4 G. W. Burton and K. U. Ingold, *Acc. Chem. Res.*, 1986, **19**, 194–201.

5 (a) P. B. Nielsen, A. Müllertz, T. Norling and H. G. Kristensen, *Int. J. Pharm.*, 2001, **222**, 217–224; (b) C.-C. Chang, J.-J. Lee, C.-W. Chiang, T. Jayakumar, G. Hsiao, C.-Y. Hsieh and J.-R. Sheu, *Pharm. Biol.*, 2010, **48**, 938–946.

6 (a) J. R. Sheu, C. R. Lee, G. Hsiao, W. C. Hung, Y. M. Lee, Y. C. Chen and M. H. Yen, *Life Sci.*, 1999, **65**, 197–206; (b) J. R. Sheu, C. R. Lee, C. C. Lin, Y. C. Kan, C. H. Lin, W. C. Hung, Y. M. Lee and M. H. Yen, *Brit. J. Pharmacol.*, 1999, **127**, 1206–1212.

7 Y. J. Suzuki and L. Packer, *Biochem. Biophys. Res. Commun.*, 1993, **193**, 277–283.

8 T. A. Thompson and G. Wilding, *Mol. Cancer Ther.*, 2003, **2**, 797–803.

9 D. Liang, J. Lin, H. B. Grossman, J. Ma, B. Wei, C. P. Dinney and X. Wu, *Cancer Causes Control*, 2008, **19**, 981–992.

10 (a) S. Witkowski and P. Wałejko, *Z. Naturforsch. B*, 2001, **56b**, 411–415; (b) S. Witkowski and P. Wałejko, *Z. Naturforsch. B*, 2002, **57**, 571–578; (c) A. Hryniewicka, P. Wałejko, J. Morzycki and S. Witkowski, *Pol. J. Chem.*, 2009, **83**, 78–80.

11 T. Parman, D. I. Bunin, H. H. Ng, J. E. McDunn, J. E. Wulff, A. Wang, R. Swezey, L. Rasay, D. G. Fairchild, I. M. Kapetanovic and C. E. Green, *Toxicol. Sci.*, 2011, **124**, 487–501.

12 (a) K. Shimoda, Y. Kondo, K. Abe, H. Hamada and H. Hamada, *Tetrahedron Lett.*, 2006, **47**, 2695–2698; (b) R. K. Uhrig, M. A. Picard, K. Beyreuther and M. Wiessler, *Carbohydr. Res.*, 2000, **325**, 72–80; (c) T. Satoh, H. Miyataka, K. Yamamoto and T. Hirano, *Chem. Pharm. Bull. (Tokyo)*, 2001, **49**, 948–953.

13 (a) S. Witkowski, D. Maciejewska and I. Wawer, *J. Chem. Soc., Perkin Trans. 2*, 2000, 1471–1476 and refs. therein; (b) S. Witkowski and I. Wawer, *J. Chem. Soc., Perkin Trans. 2*, 2002, 433–436.

14 (a) I. Wawer and S. Witkowski, *Curr. Org. Chem.,* 2001, **5**, 987–999; (b) S. Witkowski, K. Paradowska and I. Wawer, *Magn. Reson. Chem.,* 2004, **42**, 863–869; (c) D. K. Stępień, M. K. Cyrański, Ł. Dobrzycki, P. Wałejko, A. Baj, S. Witkowski, K. Paradowska and I. Wawer, *J. Mol. Struct.,* 2014, **1076**, 512–517.

15 M. Górecki, A. Suszczyńska, M. Woźnica, A. Baj, M. Wolniak, M. K. Cyrański, S. Witkowski and J. Frelek, *Org. Biomol. Chem*., 2014, **12**, 2235–2254.

16 The numbering of carbon atoms and the nomenclature as proposed by the IUPAC-IUB Joint Commission on Biochemical Nomenclature for tocopherol and related compounds,[17] was used. In the sugar part of both examined molecules, the numbering with primes (1',2', 3' *etc.*) was applied thorough the work.

17 See, *e.g.,* (a) *Eur. J. Biochem.*, 1982, **123**, 473–475; (b) *Pure Appl. Chem.,* 1982, **54**, 1507–1510.

18 K. Brzezinski, P. Wałejko, A. Baj, S. Witkowski and Z. Dauter, *Acta Crystallogr. Sect. E: Struct. Rep. Online*, 2011, **67**, o718–o718.

19 J. Tomasi, B. Mennucci and E. Cancès, *J. Mol. Struct. (Theochem),* 1999, **464**, 211–226 and refs. therein.

20 (a) J. Tomasi, B. Mennucci and R. Cammi, *Chem. Rev.,* 2005, **105**, 2999–3094. (b) G. Scalmani, M. J. Frisch, *J. Chem. Phys*., 2010, **132**, 114110/1-15.

21 (a) J.-Y. Tao, W.-H. Mu, G. A. Chass, T.-H. Tang and D.-C. Fang, *Int. J. Quantum. Chem.,* 2013, **113**, 975–984; (b) D.-C. Fang, *IDSCRF-RADII*, Beijing Normal University, Beijing, China.

22 For some most recent papers, see: (a) Y.-M. Xing, L. Zhang and D.-C. Fang, *Organometallics,* 2015, **34**, 770−777; (b) L. Zhao, and D.-C. Fang, *Eur. J. Org. Chem.,* 2015, 4772–4781; (c) L. Zhang and D.-C. Fang, *Org. Biomol. Chem*., 2015, **13**, 7950–7960; (d) W.-H. Mu, S.-Y. Xia, Y. Li, D.-C. Fang, G. Wei and G. A. Chass, *J. Org. Chem.,* 2015, **80**, 9108–9117.

23 (a) E. A. Skorupska, R. B. Nazarski, M. Ciechańska, A. Jóźwiak and A. Kłys, *Tetrahedron*, 2013, **69**, 8147–8154; (b) R. B. Nazarski, *J. Incl. Phenom. Macrocycl. Chem.,* 2014, **78**, 299–310.

24 See refs. 45a-45k cited in ref. 1c.

25 See, *e.g.,* (a) J. Vondrášek, L. Bendová, V. Klusák and P. Hobza, *J. Am. Chem. Soc.,* 2005, **127**, 2615–2619; (b) T. van Mourik, *J. Chem. Theory Comput.,* 2008, **4***,* 1610–1619; (c) S. Grimme, J. Antony, S. Ehrlich and H. Krieg, *J. Chem. Phys*., 2010, **132**, 154104/1−19; (d) S. Grimme, *Wiley Interdiscip. Rev., Comput. Mol. Sci.,* 2011, **1**, 211–228; (e) S. Grimme, S. Ehrlich and L. Goerigk, *J. Comput. Chem.*, 2011, **32**, 1456–1465; (f) N. Marom, A. Tkatchenko, M. Rossi, V. V. Gobre, O. Hod, M. Scheffler and L. Kronik, *J. Chem. Theory Comput.,* 2011, **7**, 3944–3951; (g) J. Klimeš and A. Michaelides, *J. Chem. Phys.,* 2012, **137**, 120901/1−12; (h) D. Roy, M. Marianski, N. T. Maitra and J. J. Dannenberg, *J. Chem. Phys.,* 2012, **137**, 134109/1−12; (i) S. Grimme and M. Steinmetz, *Phys. Chem. Chem. Phys.,* 2013, **15**, 16031–16042; (j) L. Goerigk, *J. Chem. Theory Comput.,* 2014, **10**, 968–980.

26 (a) L. I. Smith, H. E. Ungnade, H. H. Hoehn and S. Wawzonek, *J. Org. Chem.,* 1939, **4**, 311–317; (b). Y. Yamamoto and K. Itonaga, *Org. Lett.*, 2009, **11**, 717–720.

27 B. Helferich and E. Schmitz-Hillebrecht, *Berichte*, 1933, **66**, 378–383.

28 J. Herzig, A. Nudelman, H. E. Gottlieb and B. Fischer, *J. Org. Chem.*, 1986, **51**, 727–730.

29 R. B. Nazarski, *Magn. Reson. Chem.,* 2003, **41**, 70–74.

30 (a) J. F. Stoddart, *Stereochemistry of Carbohydrates*, Wiley-Interscience, New York, 1971, pp. 129–145; (b) D. R. Bundle and R. U. Lemieux, *Meth. Carbohydr. Chem.,* 1976, **7**, 79–86.

31 See *e.g.,* (a) J. C. Lindon and A. G. Ferrige, *Prog. Nucl. Magn. Reson. Spectrosc.*, 1980, **14**, 27–66; (b) T. Kupka and J. O. Dzięgielewski, *Magn. Reson. Chem.,* 1988, **26**, 353–357; (c) F. A. Anet and D. J. O'Leary, *Tetrahedron Lett.,* 1989, **30**, 2755–2758; (d) L. Griffiths, *Magn. Reson. Chem.,* 2001, **39**, 194–202.

32 C. Morat, F. R. Taravel and M. R. Vignon, *Magn. Reson. Chem*., 1988, **26**, 264–270.

33 See *e.g.,* papers with the 400-MHz NMR spectra of pure **2α**: (a) R. M. van Well, K. P. Ravindranathan Kartha and R. A. Field, *J. Carbohydr. Chem.,* 2005, **24**, 463–474; (b) L. Shi, G. Zhang and F. Pan, *Tetrahedron*, 2008, **64**, 2572–2575; (c) A. Kumar, Y. Geng and R. R. Schmidt, *Eur. J. Org. Chem.,* 2012, 6846–6851; (d) D. Chatterjee, A. Paul, R. Yadav and S. Yadav, *RSC Adv.,* 2015, **5**, 29669–29674.

34 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A. Petersson, H. Nakatsuji, M. Caricato, X. Li, H. P. Hratchian, A. F. Izmaylov, J. Bloino, G. Zheng, J. L. Sonnenberg, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, J. A. Montgomery, Jr., J. E. Peralta, F. Ogliaro, M. Bearpark, J. J. Heyd, E. Brothers, K. N. Kudin, V. N. Staroverov, T. Keith, R. Kobayashi, J. Normand, K. Raghavachari, A. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, N. Rega, J. M. Millam, M.

Klene, J. E. Knox, J. B. Cross, V. Bakken, C. Adamo, J. Jaramillo, R. Gomperts, R. E. Stratmann, O. Yazyev, A. J. Austin, R. Cammi, C. Pomelli, J. W. Ochterski, R. L. Martin, K. Morokuma, V. G. Zakrzewski, G. A. Voth, P. Salvador, J. J. Dannenberg, S. Dapprich, A. D. Daniels, O. Farkas, J. B. Foresman, J. V. Ortiz, J. Cioslowski and D. J. Fox, *Gaussian 09, Revision D.01*, Gaussian, Inc., Wallingford CT 06492, USA, 24-Apr-2013.

35 (a) C. Adamo and V. Barone, *J. Chem. Phys.*, 1998, **108**, 664−675; (b) M. W. Lodewyk and D. J. Tantillo, *J. Nat. Prod.,* 2011, **74**, 1339–1343; (c) M. W. Lodewyk, C. Soldi, P. B. Jones, M. M. Olmstead, J. Rita, J. T. Shaw and D. J. Tantillo, *J. Am. Chem. Soc.,* 2012, **134**, 18550−18553.

36 (a) C. Lee, W. Yang and R. G. Parr, *Phys. Rev. B*, 1988, **37**, 785–789; (b) A. D. Becke, *Phys. Rev. A*, 1988, **38**, 3098–3100; (c) A. D. Becke, *J. Chem. Phys.,* 1993, **98**, 5648–5652; (d) P. J. Stephenes, F. J. Devlin, C. F. Chabalowski and M. J. Frisch, *J. Phys. Chem.*, 1994, **98**, 11623–11627; (e) M. J. Frisch, J. A. Pople and J. S. Binkley, *J. Chem. Phys.,* 1984, **80**, 3265–3269; (f) B. J. Lynch, Y. Zhao and D. G. Truhlar, *J. Phys. Chem. A*, 2003, **107**, 1384–1388; (g) G. I. Csonka, A. D. French, G. P. Johnson and C. A. Stortz, *J. Chem. Theory Comput.,* 2009, **5***,* 679–692.

37 R. B. Nazarski, *J. Phys. Org. Chem.,* 2007, **20** 422–430.

38 (a) I. Alkorta and J. Elguero, *New. J. Chem.*, 1998, **22**, 381–385; (b) A. Perczel and A. G. Császár, *Eur. Phys. J. D,* 2002, **20**, 513–530 and their previous papers; (c) A. Buczek, M. Makowski, M. Jewgiński, R. Latajka, T. Kupka and M. A. Broda, *Biopolymers*, 2014, **101**, 28–40; (d) R. Wałęsa, T. Ptak, D. Siodłak, T. Kupka, M. A. Broda, *Magn. Reson. Chem*., 2014, **52**, 298–305; (e) R. Wałęsa, T. Kupka and M. A. Broda, *Struct. Chem.*, 2015, **26**, 1083–1093.

39 See *e.g.,* refs. 2b,4e-h,j,12,15e,18d,e and 34c cited in ref. 1c.

40 (a) *Eur. J. Biochem.,* 1983, **131**, 5–7; (b) *Pure Appl. Chem.,* 1983, **55**, 1269–1272.

41 Nomenclature of carbohydrates (Recommendations 1996) *Prepared for publication by* A D. McNaught, *Pure Appl. Chem.,* 1996, **68**, 1919–2008.

42 P. Wałejko and M. K. Cyrański, in preparation.

43 (a) V. S. Rao and A. S. Perlin, *Can. J. Chem.,* 1983, **61**, 2688–2694; (b) H. Ohrui, Y. Nishida, H. Itoh and H. Meguro, *J. Org. Chem.,* 1991, **56**, 1726–1731; (c) I. Tvaroška and J. Gajdoš, *Carbohydr. Res.,* 1995, **271**, 151–162; (d) G. D. Rockwell, T. B. Grindley and J.-P. Lepoittevin, *J. Carbohydr. Chem.,* 1999, **18**, 51–56; (e) M. Appell, G. Strati, J. L. Willett and F. A. Momany, *Carbohydr. Res.,* 2004, **339**, 537–551.

44 (a) *PCMODEL V 8.5, Molecular Modeling Software for Windows Operating System, Apple Macintosh OS, Linux and Unix*, Serena Software, Bloomington, IN 47402-3076, USA, August 2003.

45 A. Hocquet and M. Langgård, *J. Mol. Model.,* 1998, **4**, 94–112.

46 *HyperChem–Molecular Modeling System*, Release 8.0.10 for Windows, Hypercube, Inc., 1115 NW, 4th Street, Gainesville, FL 32601, USA, 2011.

47 It was taking a shortcut to reduce the amount of steps of an available GMMX searching protocol needed to find all energy minima, because 'the number of the conformations found is proportional to the time spent conducting the search'.[48]

48 W. Schepens, PhD thesis, Ghent University, Belgium, 2000 (as was cited in L. Piela*, Ideas of Quantum Chemistry*, Elsevier, Amsterdam, 2007, p. 292).

49 (a) A. F. C. Alcântara, D. Piló-Veloso, W. B. De Almeida, C. R. A. Maltha and L. C. A. Barbosa, *J. Mol. Struct.,* 2006, **791**, 180–185; (b) A. M. Belostotskii, *J. Org. Chem.,* 2008, **73**, 5723–5731; (c) G. P. Souza, C. Konzen, T. R. G. Simões, B. L. Rodrigues, A. F. C. Alcântara and H. O. Stumpf, *J. Mol. Struct.,* 2012, **1016**, 13–21.

50 (a) L. A. Nafie and T. B. Feedman, Biological and Pharmaceutical Applications of Vibrational Optical Activity in *Infrared and Raman Spectroscopy of Biological Materials*, eds. H.-U. Gremlich and B. Yan, Marcel Dekker, Inc., New York, 2001, pp. 15–54; (b) Y. He, B. Wang, R. K. Dukor and L. A. Nafie, *Appl. Spectrosc.*, 2011, **65**, 699–723; (c) L. A. Nafie, personal communication (August 14, 2015).

51 (a) L. Piela, J. Kostrowicki and H. A. Scheraga, *J. Phys. Chem.*, 1989, **93**, 3339–3346; (b) J. Kostrowicki, L. Piela, B. J. Cherayil and H. A. Scheraga, *J. Phys. Chem.,* 1991, **95***,* 4113–4119; (c) S. I. Sukharev, W. J. Sigurdson, C. Kung and F. Sachs, *J. Gen. Physiol.,* 1999, **113**, 525–539; (d) Z. Xiang, C.S. Soto and B. Honig, *Proc. Natl. Acad. Sci. USA*, 2002**, 99**,7432–7437.

52 (a) Y. Zhao and D. G. Truhlar**,** *Theor. Chem. Acc.,* 2008, **120**, 215–241; (b) Y. Zhao and D. G. Truhlar**,** *J. Phys. Chem. C*, 2008, **112***,* 4061–4067.

53 J.-D. Chai and M. Head-Gordon, *Phys. Chem. Chem. Phys.,* 2008, **10**, 6615–6620.

54 J. W. Ochterski, *Thermochemistry in Gaussian*; http://www.gaussian.com/g_whitepap/thermo/thermo.pdf; accessed August 10, 2015.

55 (a) J. Echeverría, G. Aullón, D. Danovich, S. Shaik and S. Alvarez. *Nat. Chem.,* 2011, **3**, 323−330; (b) D. Danovich, S. Shaik, F. Neese, J. Echeverría, G. Aullón and S. Alvarez, *J. Chem. Theory Comput.*, 2013, **9**, 1977−1991; (c) W. C. McKee and P. v. R. Schleyer, *J. Am. Chem. Soc.,* 2013, **135**, 13008−13014 and refs. therein.

56 Initial code names xxxxxx of all of conformers of **1α** and **1β** originate from the MMX energies (xx.xxx in kcal mol^{-1}) of their simple MMX models, after omission of the decimal point.

57 (a) D. Ŕeha, H. Valdés, J. Vondrášek, P. Hobza, A. Abu-Riziq, B. Crews and M. S. de Vries, *Chem. Eur. J.,* 2005, **11**, 6803–6817; (b) L. F. Holroyd and T. van Mourik, *Chem. Phys. Lett.,* 2007, **442**, 42–46; (c) H. Valdes, K. Pluháčková, M. Pitonák, and P. Hobza, *Phys. Chem. Chem. Phys.,* 2008, **10**, 2747–2757; (d) H. Valdes, V. Spivok, J. Řezáč, D. Reha, A. Abu-Riziq, M. S. de Vries and P. Hobza, *Chem. Eur. J.,* 2008, **14**, 4886–4898.

58 L. Ding, N. Ishida, M. Murakami and K. Morokuma, *J. Am. Chem. Soc.,* 2014, **136**, 169–178.

59 Y. Li and D.-C. Fang, *Phys. Chem. Chem. Phys.,* 2014, **16**, 15224–15230.

60 W.-H. Mu, G. A. Chass and D.-C. Fang, *Int. J. Quantum Chem.,* 2008, **108**, 1422–1434 and refs. therein.

61 (a) J.-Y. Tao, W.-H. Mu, G. A. Chass, T.-H. Tang and D.-C. Fang, *Int. J. Quantum. Chem.,* 2013, **113**, 975–984; (b) D.-C. Fang, *IDSCRF-RADII*, Beijing Normal University, Beijing, China.

62 For some most recent papers, see: (a) Y.-M. Xing, L. Zhang and D.-C. Fang, *Organometallics,* 2015, **34**, 770−777; (b) L. Zhao and D.-C. Fang, *Eur. J. Org. Chem.,* 2015, 4772–4781; (c) L. Zhang and D.-C. Fang, *Org. Biomol. Chem*., 2015, **13**, 7950–7960; (d) W.-H. Mu, S.-Y. Xia, Y. Li, D.-C. Fang, G. Wei and G. A. Chass, *J. Org. Chem.,* 2015, **80**, 9108–9117.

63 (a) P. Tähtinen, A. Bagno, K. D. Klika and K. Pihlaja, *J. Am. Chem. Soc*., 2003, **125**, 4609–4618; (b) J. N. Woodford and G. S. Harbison, *J. Chem. Theory Comput.,* 2006, **2**, 1464–1475; (c) Z. Zhang, Q.-s. Li, Y. Xie, R. B. King and H. F. Schaefer III, *New. J. Chem.,* 2010, **34**, 92–102; (d) A. Buczek, T. Kupka and M. A. Broda, *J. Mol. Model.,* 2011, **17**, 2029–2040; (e) M. A. Broda, A. Buczek and T. Kupka, *Vib. Spectrosc.*, 2012, **63**, 432–439; (f) T. Kupka, M. Stachów, E. Chełmecka, K. Pasterny, M. Stobińska, L. Stobiński and J. Kaminský, *J. Chem. Theor. Comput.*, 2013, **9**, 4275−4286.

64 Very similar approach was presented in ref. 63c.

65 D. C. Young, *Computational Chemistry: A Practical Guide for Applying Techniques to Real-World Problems*, John Wiley & Sons, New York, 2001, p. 94.

66 And being very similar to those reported[43a] for methyl 2,3,4,6-tetra-*O*-acetyl-D-glucopyranoside in $CDCl_3$ (**2α**: $^3J_{H5,H6S}$ = 2.4 and $^3J_{H5,H6R}$ = 4.6 Hz, **2β**: $^3J_{H5,H6S}$ = 2.5 and $^3J_{H5,H6R}$ = 4.6 Hz).

67 (a) F. Neese, *ORCA – An ab initio, DFT and semiempirical SCF-MO package.* Version 2.8.0.1; University of Bonn, Bonn, Germany, December 12, 2010; http://www.thch.uni-bonn.de/tc/orca; (b) F. Neese, *Wiley Interdiscip. Rev.: Comput. Mol. Sci.,* 2012, **2**, 73–78.

68 S. Sofue, T. Yamasaki, H. Morita and Y. Kitahama, *Polym. J.,* 1998, **30**, 891–896.

69 C. Møller and M. S. Plesset, *Phys. Rev.,* **1934**, *46*, 618–622.

70 Individual zero-point corrected total electronic-nuclear energies, $E_{0,i}$s, were approximated here as $E_{0,i} \cong E_{\text{tot},i}$ + $\text{ZPVE}_{\text{DFT},i}$ where $\text{ZPVE}_{\text{DFT},i}$ is an unscaled DFT-level zero-point vibrational energy correction (see Computational). Undoubtedly, the use of ΔE_0s (0 K) instead of ΔGs (288.15 K) introduced additional errors in Boltzmann populations.

71 The problem of MP2 calculations is the inaccurate description (overestimation in most cases) of long-range LD forces when an extended basis set is used. Reliable energies are thus frequently obtained with the medium size basis sets. However, this is evidently due to a compensation of errors.[57c]

72 For limitations of MP2 methods in the description of conformational energies, see also: (a) M. Goldey, A. Dutoi and M. Head-Gordon, *Phys. Chem. Chem. Phys.,* 2013, **15**, 15869–15875; (b) Y. K. Kang and H. S. Park, *Chem. Phys. Lett.,* 2014, **600**, 112–117.

73 P. Schreiner, P. v. R. Schleyer and H. F. Schaefer III, *J. Org. Chem.,* 1997, **62**, 4216–4228.

74 Such calculations were suggested by one of the referees.

75 K. B. Wiberg, *J. Comput. Chem.,* 1999, **20**, 1299–1303.

76 It should be noted that the MP2/6-311+G(*d*,*p*) optimized geometries were used in ref. 75 and that the GIAO predictions (MP2 *vs.* DFT-mPW1PW91) of σ_Cs were performed applying the 6-311+G(*2d*,*p*) basis set. Moreover, the GIAO/MP2 procedure was not as successful; mPW1PW91 gave smaller errors.[75] The lack of diffuse functions was most likely critical for our GIAO/MP2 results, but their use was computationally prohibitively expensive.

77 It is obvious that adequate scripting and data-management techniques used as a screening tool in the initial step of this and similar studies (as suggested by one reviewer) would make such endeavors manageable for many other multi-conformer systems.

78 See also (a) E. Kolehmainen, K. Laihia, M. Heinänen, K. Rissanen, R. Fröhlich, J. Korvola, P. Mänttäri and R. Kauppinen. *J. Chem. Soc., Perkin Trans. 2*, 1993, 641–648; (b) Ref. 22 cited in M. H. Chisholm, N. W. Eilerts, J. C. Huffman, S. S. Iyer, M. Pacold and K. Phomphrai, *J. Am. Chem. Soc.*, 2000, **122**, 11845–11854.

79 See, *e.g.,* (a) M. Saunders, *J. Am. Chem. Soc*., 1987, **109**, 3150–3152; (b) M. Saunders, *J. Comp. Chem.,* 1989, **10**, 203–208; (c) M. Saunders, K. N. Houk, Y.-D. Wu, W. C. Still, M. Lipton, G. Chang and W. C. Guida, *J. Am. Chem. Soc*., 1990, **112**, 1419–1427.

80 (a) K. S. Steliou, *BAKMDL*, 1989, based on the original program by W. C. Still; (b) Y. Murakami, J.-i. Kikuchi, T. Ohno, T. Hirayama, Y. Hisaeda, H. Nishimura, J. P. Snyder and K. Steliou, *J. Am. Chem. Soc*., 1991, **113**, 8229–8242 and refs therein; (c) M. M. Midland, G. Asirwatham, J. C. Cheng, J. A. Miller and L. A. Morell, *J. Org. Chem.*, 1994, **59**, 4438–4442 and refs therein.
81 D. M. Ferguson and D. J. Raber, *J. Am. Chem. Soc*., 1989, **111**, 4379–4386.
82 (a) S. R.Wilson, W. Cui, J. W. Moskowitz and K. E. Schmidt, *Tetrahedron Lett.,* 1988, **29**, 4373-4376; (b) G. Chang, W. C. Guida and W. C. Still, *J. Am. Chem. Soc*., 1989, **111**, 4371–4378.
83 J. J. Gajewski, K. E. Gillbert and J. McKelvey, MMX: An enhanced version of MM2 in *Advances in Molecular Modeling*, ed. D. Liotta, JAI Press, Inc., London, 1990, vol. 2, pp. 65–92.
84 D. Jiao, M. Barfield and V. J. Hruby, *J. Am. Chem. Soc*., 1993, **115**, 10883–10887.
85 W. Migda and B. Rys, *Magn. Reson. Chem.,* 2004, **42**, 459–466 and refs. 10-15 cited therein.
86 J. Gräfenstein and D. Cremer, *J. Chem. Phys.,* 2007, **127**, 164113/1–7.
87 *ChemCraft,* Version 1.7 (built 375); http://www.chemcraftprog.com.
88 (a) R. B. Nazarski, unpublished results (2013); (b) A. Marciniak, Master's Thesis, University of Łódź, Faculty of Chemistry, Łódź, 2014; (c) R. B. Nazarski and P. Wałejko, VIIIth Symposium on: *Nuclear Magnetic Resonance in Chemistry, Physics and Biological Sciences*, Warsaw, 24-26 September, 2014, poster; Programme & Book of Abstracts, page P-21.
89 See *e.g.,* (a) K. W. Wiitala, T. R. Hoye and C. J. Cramer, *J. Chem. Theory Comput.,* 2006, **2**, 1085–1092; (b) K. Dybiec and A. Gryff-Keller, *Magn. Reson. Chem*., 2009, **47**, 63–66; (c) R. F. Ribeiro, A. V. Marenich, C. J. Cramer and D. G. Truhlar, *J. Phys. Chem. B*, 2011, **115**, 14556–14562.
90 (a) J. B. Foresman and Æ. Frisch, *Exploring Chemistry with Electronic Structure Methods*, 2nd edn., Gaussian, Inc., Pittsburgh, PA 15106, USA, 1996, Chapter 4 + Errata; (b) R. Jain, T. Bally and P. R. Rablen, *J. Org. Chem.,* 2009, **74**, 4017–4023.
91 See *e.g.,* (a) H. Dodziuk, M. Jaszuński and W. Schilf, *Magn. Reson. Chem.,* 2005, **43**, 639–646; (b) S. Leśniak, A. Chrostowska, D. Kuc, M. Maciejczyk, S. Khayar, R. B. Nazarski and Ł. Urbaniak, *Tetrahedron*, 2009, **65**, 10581–10589; (c) R. B. Nazarski and W. Makulski, *Phys. Chem. Chem. Phys.,* 2014, **16**, 15699–15708.
92 (a) S. Huzinaga, *J. Chem. Phys.,* 1965, **42**, 1293–1302; (b) S. Huzinaga, *Approximate Atomic Functions. Technical Report,* University of Alberta, Edmonton, Canada, 1971; (c) M. Schindler and W. Kutzelnigg, *J. Chem. Phys.,* **1982**, *76*, 1919–1933; (d) W. Kutzelnigg, U. Fleischer and M. Schindler, The IGLO-Method: Ab-initio Calculation and Interpretation of NMR Chemical Shifts and Magnetic Susceptibilities in *NMR - Basic Principles and Progress (Deuterium and Shift Calculation)*, eds. P. Diehl, E. Fluck, H. Günther, R. Kosfeld and J. Seelig, Springer-Verlag, Berlin, 1991, vol. 23, pp. 165–262.
93 See refs. 38-43 cited in ref. 91c.
94 (a) EMSL Basis Set Library; https://bse.pnl.gov/bse/portal; (b) D. Feller, *J. Comp. Chem.,* 1996, **17**, 1571–1586; (c) K. L. Schuchardt, B. T. Didier, T. Elsethagen, L. Sun, V. Gurumoorthi, J. Chase, J. Li, and T. L. Windus, *J. Chem. Inf. Model.,* 2007, **47**, 1045–1052.
95 S. M. Koskowich, W. C. Johnson, R. S. Paley and P. R. Rablen, *J. Org. Chem*., 2008, **73**, 3492–3496.
96 T. Chai and R. R. Draxler, *Geosci. Model Dev.*, 2014, **7**, 1247–1250.
97 S. G. Smith and J. M. Goodman, *J. Org. Chem.,* 2009, **74**, 4597–4607.