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PAPER

Surface modification with thermoresponsive polymer brushes for a switchable electrochemical sensor

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Elaboration of switchable surfaces represents an interesting way for the development of a new generation of electrochemical sensors. In this paper, a method for growing thermoresponsive polymer brushes from a gold surface pre-modified with polyethyleneimine (PEI), subsequent layer-by-layer polyelectrolyte assembly and adsorption of a charged macroinitiator is described. We propose an easy method for monitoring the coil-to-globule phase transition of the polymer brush using electrochemical quartz crystal microbalance with dissipation (E-QCM-D). The surface of these polymer modified electrodes shows reversible switching from the swollen to the collapsed state with temperature. As demonstrated from E-QCM-D measurements using an original signal processing method, the switch is operating in three reversible steps related to different interfacial viscosities. Moreover, it is shown that the one electron oxidation of ferrocene carboxylic acid is dramatically affected by the change from the swollen to the collapsed state of the polymer brush, showing a spectacular 86 % decrease of the charge transfer resistance between the two states.

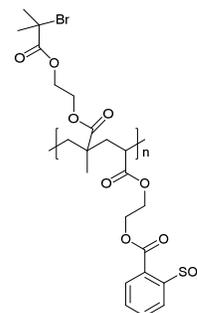
Introduction

Stimuli responsive polymers are macromolecules that can undergo structural changes as a response to external stimuli. The most common triggers for this transition are changes in temperature, pH or ionic strength¹. When such polymers are grafted on a substrate as dense brush, interfacial properties such as surface adhesion, wettability or interfacial viscoelastic properties can be reversibly modified as result of transition induced structural changes²⁻⁴. This makes such responsive brushes attractive candidates as switchable elements in electrochemical systems^{5,6}.

In the case of thermoresponsive polymers in aqueous media, the coil-to-globule transition is mostly characterized by a lower critical solution temperature (LCST). Below the transition temperature, the grafted polymer chains are strongly swollen and stretch away from the surface, whereas above the transition, such brushes expel most of the solvent, deswell and finally collapse on the substrate. The transition temperature is characteristic for the polymer studied and the specific aqueous solvent used. It can be conveniently tuned by co-polymerization of different monomers. Macromolecules constructed with short oligo(ethylene glycol) methacrylates (OEGMA) constitute an interesting class of thermoresponsive polymers for preparing biorelevant and biocompatible switchable surfaces⁷. While the phase transition of these copolymers is reversible, the precise transition temperature of a given copolymer depends to some extent on system parameters such as molar mass and polydispersity, polymer concentration, and the ionic strength of the medium⁷⁻¹¹. Random

copolymers of 2-(2-methoxyethoxy)ethyl methacrylate (MEO₂MA) and OEGMA typically exhibit a LCST in water, which can be precisely adjusted by varying the co-monomer ratio^{7, 10, 12-15}. Interestingly, this LCST-type transition can be detected by electrochemical methods¹⁶.

We herein report a molecular architecture of a LCST polymer modified gold surface, and its interfacial properties in terms of charge transfer towards an electrochemical probe. The modular character of the straightforward coating process enables a priori the simple and versatile functionalization of such polymer modified electrodes. We also introduce differential E-QCM-D measurement to characterize the ability of the system to actively respond upon temperature variations in order to develop a switchable readout system exploiting the structural changes of the surface.



Scheme 1 Structure of the macroinitiator (MI)

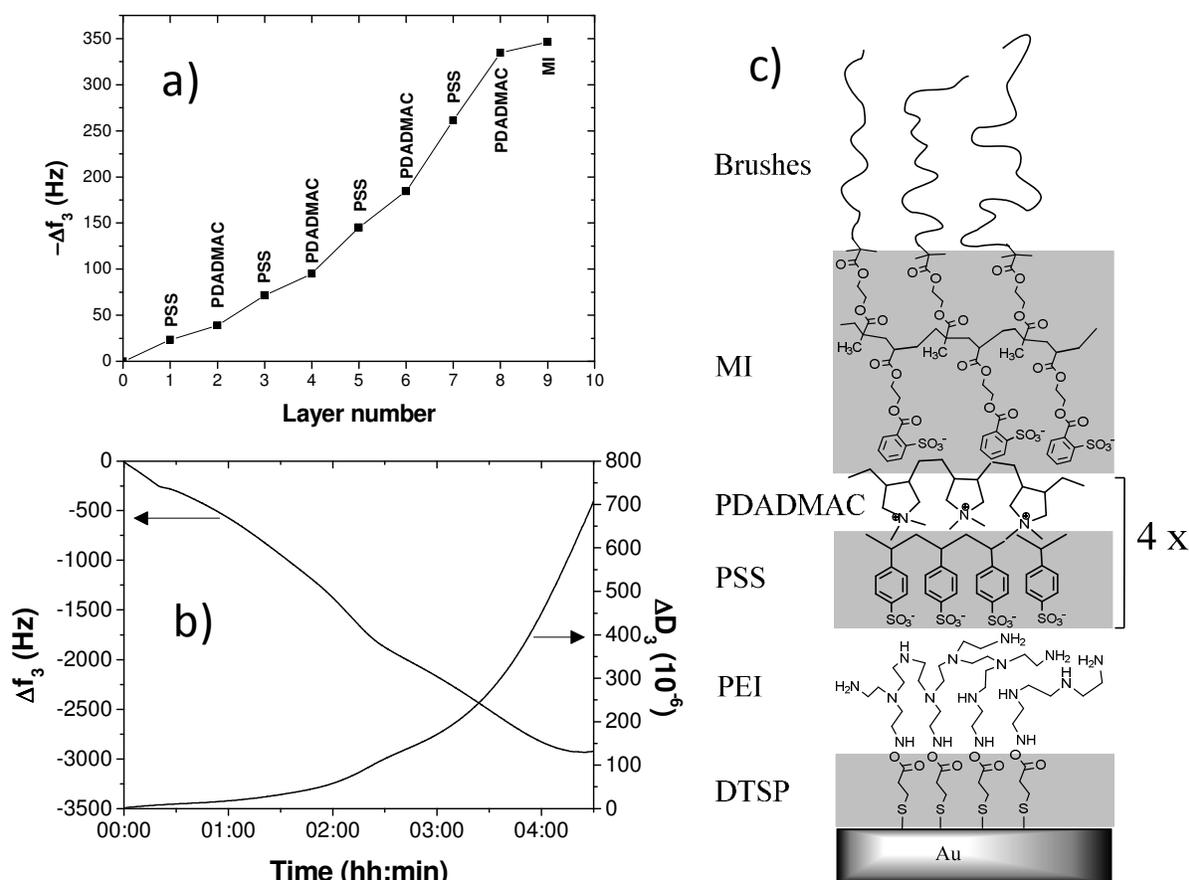


Fig.1 Resonance frequency and dissipation shift monitored by E-QCM-D upon (a) LBL assembly by successive deposition of PSS and PDADMAC and finally of the macroinitiator MI, and (b) polymerization from the pre-modified electrode surface. (c) Proposed structure of the smart PME (scheme)

Experimental methods

Chemicals

Poly(4-styrenesulfonic acid) (PSS, $M_w = 70,000 \text{ g}\cdot\text{mol}^{-1}$), Ferrocene carboxylic acid (FcCOOH), Tris(hydroxymethyl)aminomethane (Tris), hydroxyethyl methacrylate (HEMA), Dithiobis(succinimidyl propionate) (DTSP), branched polyethyleneimine (PEI, $M_w = 25,000 \text{ g}\cdot\text{mol}^{-1}$), and azo-bis-isobutyronitrile (V-601) were purchased from Sigma (Steinheim, Germany) and used as received, when not specified otherwise. Poly(diallyldimethylammonium chloride) (PDADMAC, $M_n = 230,000 \text{ g}\cdot\text{mol}^{-1}$, $M_w = 570,000 \text{ g}\cdot\text{mol}^{-1}$) was prepared according to a published procedure¹⁷.

Surface modification: DTSP self assembled monolayer (SAM), coupling with PEI and layer by layer assembly

All surface modifications were monitored with an Electrochemical Quartz Crystal Microbalance with Dissipation (E-QCM-D), as it is a well suited method for monitoring the consecutive surface modification steps¹⁸, using a Q-Sense E4-instrument (Q-Sense, AB, Göteborg, Sweden) incorporating a Q-

Sense Electrochemistry Module, QEM 401 and polished AT-cut piezoelectric quartz crystals with gold electrodes with a fundamental frequency of 5 MHz as substrates. The gold substrates (disk for E-QCM-D and wires for electrochemical measurements) were cleaned according to a previously published procedure¹⁹. Briefly, crystals were immersed in a mixture of ammonia (25%), hydrogen peroxide (30%), and Milli-Q water (1:1:5), held at 70 °C for 10 min, rinsed extensively with Milli-Q water, and dried under a stream of nitrogen.

A first modifying layer consisting of a SAM of the functional disulfide DTSP was prepared by incubating the clean gold electrode in a $10 \text{ mmol}\cdot\text{l}^{-1}$ DTSP solution in CHCl_3 at 4 °C overnight. DTSP modified electrodes were rinsed with CHCl_3 and water and assembled in the E-QCM-D module. PEI aqueous solution ($1 \text{ g}\cdot\text{l}^{-1}$) was flown ($10 \mu\text{l}\cdot\text{min}^{-1}$) over the SAM to covalently couple the primary amine moieties of PEI to DTSP. Then, layer-by-layer assembly (LBL)^{20, 21} of 4 alternate layer pairs of PSS and PDADMAC ($0.2 \text{ g}\cdot\text{l}^{-1}$ solution in $5 \text{ mmol}\cdot\text{l}^{-1}$ Tris buffer pH 7.2) was applied in order to build a polyelectrolyte multilayer. Each deposition step was followed by rinsing with $5 \text{ mmol}\cdot\text{l}^{-1}$ Tris buffer pH 7.2 to remove the loosely adhering

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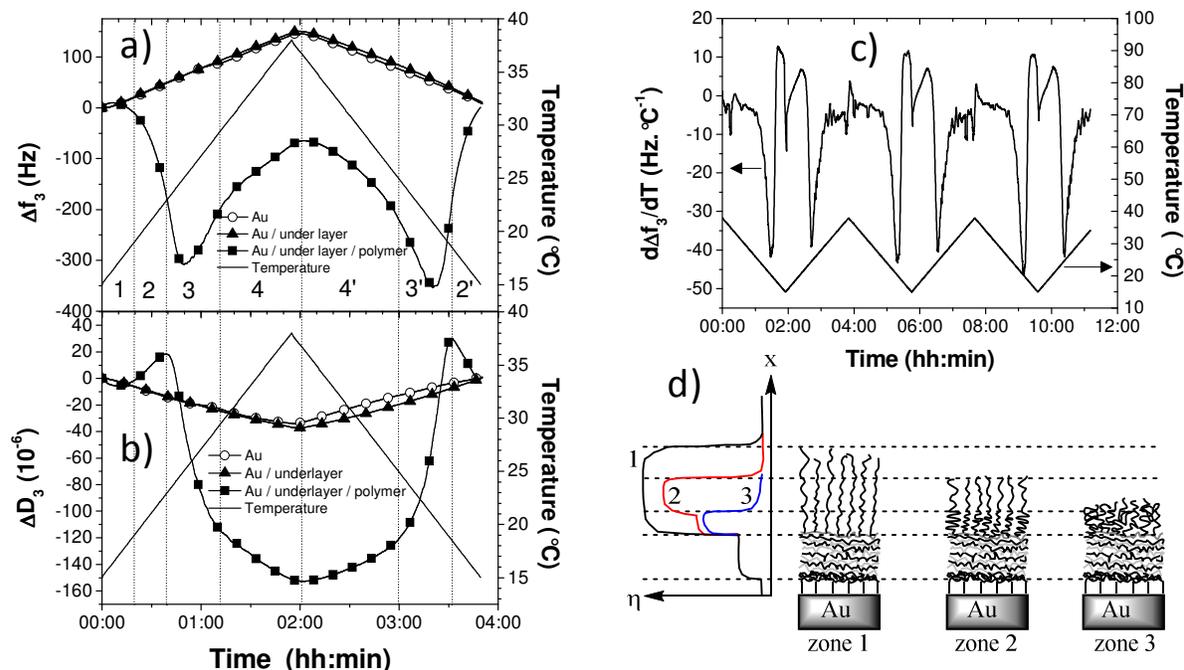


Fig.2 (a) Resonance frequency shift and (b) dissipation shift and temperature change for Au, Au/PEI/[PSS-PDADMAC]₄/MI and PME. vs. time. (c) Derivative of the frequency over the temperature of PME for 3 thermal cycles. (d) Scheme of the three states of the polymer layer involving sequential collapse of polymer layers, and their proposed interfacial viscosity (η) evolution vs. distance from the surface (x).

5 molecules. The flow rate was always 50 $\mu\text{l}\cdot\text{min}^{-1}$, if not otherwise mentioned, and the injected solution was changed only after a stable frequency signal was reached. Finally, the sulfonated polyanion macroinitiator MI (scheme 1) (0.2 $\text{mg}\cdot\text{ml}^{-1}$ in water), synthesized according to a published procedure²², was adsorbed
10 onto the terminal PDADMAC layer.

Surface initiated polymerization

The MI on top of the polyelectrolyte multilayer was used as starting point for the surface-initiated atom transfer radical random copolymerization (ATRP) of MEO₂MA and OEGMA
15 monomers (ratio 95:5),²³⁻²⁵ which was catalysed with a copper 2,2'-bipyridine bromide complex ($\text{CuBr}_2\text{bpy}_2$). A stock solution of co-monomers and copper complex is bubbled with nitrogen and pumped into the E-QCM-D chamber. The polymerization was monitored by following both frequency shift and dissipation
20 with time. The resulting modified surface will be referred to as the Polymer Modified Electrode (PME) in the following. The phase transition temperatures of the grafted polymer brushes PME were determined by E-QCM-D measurements using a temperature ramp of 0.2 $^\circ\text{C}\cdot\text{min}^{-1}$ from 15 $^\circ\text{C}$ to 38 $^\circ\text{C}$ in a 25
25 $\text{mmol}\cdot\text{l}^{-1}$ TRIS buffer pH 8.

Electrochemical measurements

Cyclic voltammetry and electrochemical impedance spectroscopy (EIS) were performed on a CH Instrument Model 750 A (Austin,
50

USA) to characterize the electronic transfer for FcCOOH oxidation (5 $\text{mmol}\cdot\text{l}^{-1}$ solution in citrate-phosphate buffer 0.1
30 $\text{mol}\cdot\text{l}^{-1}$ pH 7.1). A thermostated 3 electrode cell comprising the polymer modified electrode (PME gold wire) as working electrode, Ag/AgCl, KCl 1M reference and platinum wire as counter electrode were used. All potentials are referred to
35 Ag/AgCl, KCl 1M reference. DC potential for EIS measurements was chosen as the potential obtained at $I_p^a/2$ during cyclic voltammetry at 10 $^\circ\text{C}$ and 50 $\text{mV}\cdot\text{s}^{-1}$ (310 mV for PME and 282 mV for the bare gold electrode). Amplitude and frequency range are 10 mV and 96 kHz to 374 mHz respectively. Charge transfer
40 resistance (R_{CT}) for the one electron oxidation of FcCOOH was determined by fitting from R_e -(R_{ct} /CPE) equivalent circuit using Z View software.

Results and discussion

Surface modification

45 The E-QCM-D measurements show that PEI covalent grafting onto the DTSP modified gold substrate is completed after 1h, with a 23 Hz negative frequency shift (supporting figure 1) (3rd overtone), which corresponds to a mass of 136 $\text{ng}\cdot\text{cm}^{-2}$ using the Sauerbrey equation (equation 1).

$$\Delta m = -\frac{c\Delta f}{n} \quad (1)$$

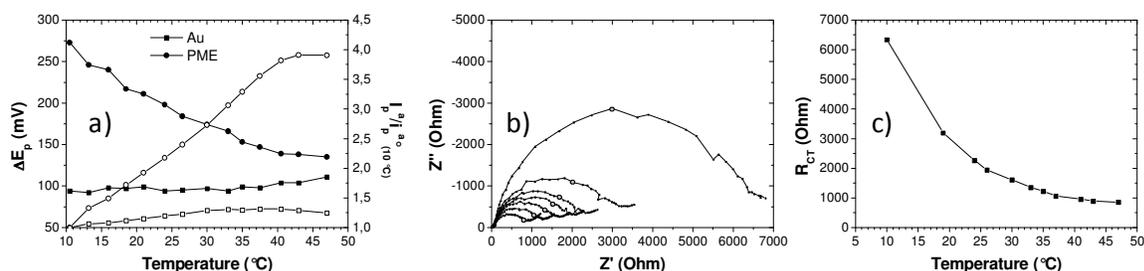


Fig.3 (a) ΔE_p (filled symbols) and r (open symbols) vs. temperature ($50 \text{ mV}\cdot\text{s}^{-1}$) on bare Au (squares) and on PME (circles) electrodes, (b) EIS: full circle corresponds to 14 Hz. From largest to smallest semi-circle: 10; 19; 24; 26; 30; 35 and 47 °C and (c) R_{CT} vs. temperature for the PME.

5 Where Δm is the mass uptake ($\text{ng}\cdot\text{cm}^{-2}$), c is a constant related to the quartz crystal ($17.7 \text{ ng}\cdot\text{Hz}^{-1}\cdot\text{cm}^{-2}$ for a 5 MHz crystal) and n is the overtone number. Most of the time, the fundamental resonance frequency f_1 is noisy because of energy trapping or piezoelectric stiffening²⁶. We therefore used the 3rd overtone in
10 all measurements.

The LBL assembly was monitored by E-QCM-D (fig. 1a). A negative frequency shift is observed after each layer deposition. A final frequency shift of 350 Hz is measured for a chip covered by 4 bilayers of PSS/PDADMAC and one layer of macroinitiator.
15 The corresponding dissipation is rather high (84×10^{-6}) (supporting figure 2) due to the flexibility of the polyelectrolyte chains, so the Sauerbrey equation was not applicable²⁷. Fig. 1b shows the on-line monitoring of the polymerization with E-QCM-D, which leads to 2700 Hz frequency decrease after 4.5 h of
20 polymerization, indicating a large mass increase over the surface and 700.10^{-6} dissipation increase which is typical for the formation of a viscoelastic film²⁸. The proposed final PME architecture is represented on Fig. 1c. According to previous studies on similar structures²³, thicknesses of LBL underlayer (\sim
25 40 nm), MI layer (\sim 20 nm) and polymer layer (\sim 70 nm) were obtained.

Phase transition of the responsive polymer

Fig. 2a and 2b show the resonance frequency and dissipation variation upon temperature changes. A bare gold chip and an assembled system without polymer layer show linear evolution of frequency and dissipation change with temperature, which are expected for solvent density and viscosity variation. However, frequency and dissipation change displays a totally different behavior for the PME. Upon temperature increase, a large
35 negative frequency shift is observed. This is related to the collapsing of the polymer, passing from a stretched and viscoelastic form to a more rigid structure. According to Laloyaux *et al.*²⁹, this can be described as a two-step collapsing. Zone 1 corresponds to the fully swollen state, where frequency and dissipation data are close to the underlayer behavior. In zone 2, dissipation is increasing whereas frequency decreases. This difference with the under layer response shows the beginning of the bulk brushes collapsing. The increase of dissipation in that zone indicates a metastable behavior of the brushes which appear
40 to be more flexible. In zone 3, the surface brushes are fully collapsed, indicated by a large dissipation and frequency decrease. The frequency shows a minimum, indicating a maximal synergy between surface and bulk collapse. After this minimum,

the frequency change is increasing again, indicating that surface collapse is progressively completed and the polymer chains are confined and almost no more flexible. In zone 4 and 4', the film is behaving like the under layer as a rigid coating on the surface. Zones 3' and 2' are the reversible processes of surface and bulk re-swelling upon cooling.

55 As the Au/under layer assembly behaves as a thick rigid film, Δf and ΔD follow the linear density and viscosity dependency on temperature. Therefore, plotting the change in frequency over the change in time (i.e. the first derivative) versus the temperature, the under layer dependency on temperature for frequency data
60 will be removed. This enables the observation of the thermoresponsive behavior independently from solvent density and viscosity changes. Such a plot is presented in fig. 2c, where three cooling / heating cycles applied to the system are shown. Clearly, the switching is fully reversible, showing mean peak frequencies at 20.5°C for cooling ramps and 24.6 °C for heating ramps. Taking the hysteresis (coming from the time constant of the system) into account, the average of these two mean values is representative of a steady state measurement, what gives a frequency peak at 22.6 °C, in good agreement with a reported
70 transition temperature of 24.6°C for brushes with a similar composition grown on SiO_2 surfaces²⁹. Fig. 2d schematically shows the two steps collapse. These structural changes create new interfaces that modulate the interfacial viscosity.

Electrochemistry of FcCOOH and barrier effect of the polymer film

We performed the electrochemical oxidation of ferrocene carboxylic acid (FcCOOH) as a function of temperature in order to investigate the barrier effect of the polymer film towards electron transfer and diffusion. Since the transition temperature
80 determined by E-QCM-D measurements is close to 22.6 °C, the temperature was varied between 10 and 47°C. Fig. 3a shows the peak separation (ΔE_p) and the ratio of anodic peak current over the reference peak current in a fully swollen state at 10 °C ($r = i_p^a/i_p^{a0}$) as a function of temperature for both a bare gold electrode
85 and PME. Comparison of the reference temperature dependency on bare gold with PME demonstrates that ΔE_p and r differ strongly for the two electrodes. For both parameters, values shift continuously in the temperature range of 10 to about 40°C, where a plateau value is reached, in the case of PME, indicating that
90 both the kinetics for electron transfer and diffusion issues are greatly affected by the polymer phase transition. ΔE_p and r are linearly evolving till reaching a plateau, corresponding to the

fully collapsed state, indicating a gradual thermal collapse. This finding is in agreement with temperature dependent thickness studies of similar surface brushes of poly(OEGMA-co-MEO₂MA) copolymers³⁰. EIS was performed at several temperatures (fig. 3b) ranging from 10 °C (fully swollen polymer) to 47 °C (fully collapsed polymer). Fig. 3c shows the R_{CT} dependency on temperature. The temperature of 22.6°C, where the collapsing was observed to be the most intense from E-QCM-D measurements, corresponds to the largest changes in R_{CT} . R_{CT} is no more strongly affected at temperatures above 30-33°C. This corresponds to the beginning of zone 4 in figs. 2a and b, where the polymer is considered to approach the maximum collapsed state. Finally, an 86 % decrease of R_{CT} is observed between fully swollen state (10 °C) and collapsed state (47 °C). This spectacular evolution of R_{CT} points out the strong barrier effect of the swollen polymer.

Conclusions

We demonstrate that thermoresponsive brushes made of poly(OEGMA-co-MEO₂MA) can be grown by surface initiated ATRP from gold substrates stepwise modified by a covalently attached polyamine layer and subsequent layer-by-layer (LBL) deposition of PSS, PDADMAC and of a polyelectrolyte macroinitiator as top layer. The modular character of the nevertheless rather simple built-up procedure allows for versatile and modular functionalizations of such polymer modified electrodes. As determined by electrochemical quartz crystal microbalance with dissipation (E-QCM-D), the film shows reversible switching between the swollen and the collapsed states by changing the temperature around the lower critical solution temperature (LCST). The dissipation profile with temperature suggests a two steps mechanism of the transition. This is confirmed by electrochemical oxidation of ferrocene carboxylic acid (FcCOOH) using the PME at different temperatures. The charge transfer resistance is gradually decreased with ongoing dehydration of the polymer brush while temperature is raised, till reaching a constant value when the polymer is fully collapsed. We believe that this large change of charge transfer resistance associated to the polymer state can be further implemented into switchable electrochemical systems.

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Notes

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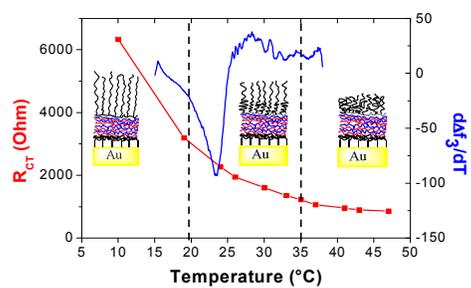
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‡ Electronic Supplementary Information (ESI) available: covalent coupling of PEI onto DTSP and dissipation shift during LBL assembly. See DOI: 10.1039/b000000x/

Table of content



Thermoresponsive polymer brushes reversible phase transition is evidenced with E-QCM-D and electrochemistry, and acts as an electrochemical switch.