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Photochemically Triggered Cheletropic Formation of Cyclopropenone (c- C_3H_2O) from Carbon Monoxide and Electronically Excited Acetylene

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ABSTRACT

For more than half a century, pericyclic reactions have played an important role in advancing our fundamental understanding of cycloadditions, sigmatropic shifts, group transfer reactions, and electrocyclicization reactions. However, the fundamental mechanisms of photochemically activated cheletropic reactions have remained contentious. Here we report on the simplest cheletropic reaction: the [2+1] addition of ground state ^{18}O -carbon monoxide (C^{18}O , $X^1\Sigma^+$) to D2-acetylene (C_2D_2) photochemically excited to the first excited triplet (T1), second excited triplet (T2), and first excited singlet state (S1) at 5 K, leading to the formation of D2- ^{18}O -cyclopropenone (c- $\text{C}_3\text{D}_2^{18}\text{O}$). Supported by quantum-chemical calculations, our investigation provides persuasive testimony on *stepwise* cheletropic reaction pathways to cyclopropenone via excited state dynamics involving the T2 (non-adiabatic) and S1 state (adiabatic) of acetylene at 5 K, while the T1 state energetically favors an intermediate structure that directly dissociates after relaxing to the ground state. The agreement between experiments in low temperature ices and the excited state calculations signifies how photolysis experiments coupled with theoretical calculations can untangle polyatomic reactions with relevance to fundamental physical organic chemistry at the molecular level, thus affording a versatile strategy to unravel exotic non-equilibrium chemistries in cyclic, aromatic organics. Distinct from traditional radical–radical pathways leading to organic molecules on ice-coated interstellar nanoparticles (interstellar grains) in cold molecular clouds and star-forming regions, the photolytic formation of cyclopropenone as presented changes the perception of how we explain the formation of complex organics in the interstellar medium eventually leading to the molecular precursors of biorelevant molecules.

1. INTRODUCTION

For more than half a century, pericyclic reactions—reactions that proceed via a single, concerted transition state allowing a cyclic overlap of π and/or σ orbitals—have played a central role in advancing our fundamental understanding of cycloadditions,¹ sigmatropic shifts,² group transfer reactions,² electrocyclicization,³ and cheletropic reactions.⁴⁻⁵ The mechanistical understanding of pericyclic reactions has impacted electronic structure and chemical bonding theory² as well as preparative organic chemistry on the industrial scale (Scheme 1).⁶⁻⁷ Early formulations of the pericyclic selection rules evolved from Frontier Molecular Orbital (FMO) theory,⁸ where a

pericyclic reaction is coined *symmetry-forbidden* or *symmetry-allowed* to indicate existence or absence of a symmetry-imposed barrier, respectively.⁹ These symmetry considerations are derived from a correlation diagram connecting the ground and excited electronic states of the reactants and products. For instance, symmetry-forbidden reactions proceed with barriers reaching up to 500 kJ mol⁻¹ for thermal [2+2] cycloadditions of two ethylene (C₂H₄) molecules to cyclobutadiene (C₄H₈).¹⁰⁻¹¹ Cheletropic reactions, where one reactant forms both bonds from the same atom with a co-reactant (Scheme 2), are pericyclic reactions that have received much attention for their utility in preparing cyclic molecules from acyclic precursors.¹²⁻¹³ The symmetry allowed [2+1] additions of electronically excited singlet carbene (CH₂, a¹A₁) to the π-electronic system of ethylene (C₂H₄) and acetylene (C₂H₂), leading to cyclopropane (C₃H₆) and cyclopropene (C₃H₄), respectively, are among the most influential cycloadditions from the fundamental and synthetic point of view.¹⁴⁻¹⁸ Cheletropic reactions have been exploited in preparative organic chemistry to synthesize, e.g., annulenes and prodrugs via carbon monoxide (CO) elimination,¹⁹ chiral phosphonium cations,²⁰ enantiomerically-pure carbocycles,²¹ and exotic enediynes.²²

Despite their importance, there is still an incomplete understanding of the fundamental mechanisms of photochemically activated cheletropic reactions. Debated mechanistic aspects are linear versus non-linear approach geometry of reactants,²³⁻²⁴ a Hückel versus Möbius (anti) aromatic character of the π-system,²⁵ the excited state surfaces and crossings involved,²⁶ and concerted versus stepwise mechanisms.²⁷ An disentangling of these key questions requires an experimental investigation of cheletropic reactions involving simple, prototype systems, which can be benchmarked and compared to computational studies. The insights derived from such ‘simple’ benchmarks are critical not only to understand the reactivity of cheletropic reactions on a molecular level, but can also indicate how to control the outcome of (stereospecific) reaction pathways in more complex systems of practical importance, e.g., for synthetic chemistry, drug design, and astrochemistry.

Here, we report on the simplest excited-state cheletropic reaction: the [2+1] photochemical addition of ground state ¹⁸O- carbon monoxide (C¹⁸O, X¹Σ⁺) to electronically excited D₂-acetylene (C₂D₂) leading to the formation of D₂-¹⁸O-cyclopropenone (c-C₃D₂¹⁸O, **3**; equation (1)).



Specifically, we investigate D₂-acetylene excited to the first excited triplet (T₁), second excited triplet (T₂), and first excited singlet state (S₁) by 222 nm (5.58 eV), 249 nm (4.98 eV), and 288 nm

(4.31 eV), respectively (Figure 1). Experiments are carried out at low temperature ices at 5 K. The photon energies employed are below the first excited electronic state of carbon monoxide (6.010 ± 0.04 eV),²⁸ the carbon-deuterium bond energy (5.797 ± 0.006 eV),²⁹ and the adiabatic ionization energies of carbon monoxide (14.0142 ± 0.0003 eV)³⁰ and D2-acetylene (11.416 ± 0.006 eV).³¹ Therefore, the photon energies used exclude ion-molecule reactions and simple carbon-deuterium bond rupture processes. These wavelengths were chosen to explore the effect of distinct low-lying electronically excited states on the low-temperature reactivity of D2-acetylene and how they influence the cheletropic character of the reaction. These surface-science experiments are merged with multi-reference *ab initio* calculations exploring the formation mechanism of D2-¹⁸O-cyclopropanone (*c*-C₃D₂¹⁸O, **3**). In addition to insight into mechanistic physical organic (photo)chemistry, an understanding of cheletropic reactions in the carbon monoxide – acetylene system is also of fundamental importance to accessing the scantily explored class of aromatic cyclic ketones (Scheme 3) with cyclopropanone (*c*-C₃H₂O, **3**) being their simplest representative. Along with its larger members cyclopentadienone (*c*-C₅H₄O, **5**) and cycloheptatrienone (*c*-C₇H₆O, **6**), cyclic π -conjugated ketones have attracted considerable attention due to their potential Hückel (anti)aromaticity,³²⁻³⁹ shedding light on the electronic structure and dimerization reactions of cyclic aromatic ketones.^{38, 40} The 2π and 6π aromaticity of cyclopropanone (**3**) and cyclopentadienone (**5**), respectively,⁴¹ is reflected in the dipolar resonance structures of the carbonyl double bond ($\text{C}=\text{O} \leftrightarrow {}^+\text{C}-\text{O}^-$) (Scheme 3)³⁸ formally leading to a carbocation with a vacant p-orbital able to foster a ring current within the delocalized 2π aromatic system.⁴²

2. RESULTS

2.1 Infrared Spectroscopy. The experiments were carried out in an ultrahigh vacuum (UHV) surface-science chamber operated at pressures of a few 10^{-11} torr by photolyzing binary ices of ¹⁸O-carbon monoxide (C¹⁸O) and D2-acetylene (C₂D₂) at photon energies of 222 nm (5.58 eV), 249 nm (4.98 eV), and 288 nm (4.31 eV) at 5 K. Infrared spectroscopy signifies an important analytical tool to identify functional groups emerging during the photoactivation (Figure S1 and Tables S1). Infrared spectra of the pristine C¹⁸O – C₂D₂ ices at 5 K revealed prominent absorptions of D2-acetylene and ¹⁸O-carbon monoxide such as the CD stretching modes (ν_1 , 2680 cm⁻¹; ν_3 , 2408 cm⁻¹),⁴³ the C¹⁸O stretching mode (ν_1 , 2088 cm⁻¹),⁴⁴ the combination modes for C₂D₂ ($\nu_1 +$

ν_3 , 5042 cm^{-1} ; $\nu_1 + \nu_5$, 3294 cm^{-1} ; $\nu_3 + \nu_4$, 2929 cm^{-1} ; $\nu_2 + \nu_5$, 2327 cm^{-1} ; $\nu_4 + \nu_5$, 1077 cm^{-1}),⁴³ and the overtone of the stretching mode of C^{18}O ($2\nu_1$, 4150 cm^{-1}).⁴⁵ At 222 nm, the photoactivation resulted in a slight decrease of $4.1 \pm 3.3\%$ and $1.3 \pm 0.8\%$ of the C_2D_2 and C^{18}O reactants monitored through their stretching mode at 2408 cm^{-1} (ν_3) and 2088 cm^{-1} (ν_1), respectively (SI Appendix, Figure S1). Simultaneously, new absorptions emerged at 1734 cm^{-1} , 1772 cm^{-1} and 2573 cm^{-1} with the 1734 cm^{-1} feature possibly linked to the carbonyl ($\text{C}=\text{O}$) functional group of D2-cyclopropanone (ν_2 of *c*- $\text{C}_3\text{D}_2^{18}\text{O}$). The remaining absorptions can be associated with carbonyl (C^{18}O) and acetylenic C-D ($-\text{C}\equiv\text{C}-\text{D}$) functional groups.^{14,38} However, due to the strong overlap of functional groups of the photoactivated mixture of organic molecules (carbonyl moieties in particular),⁴⁵⁻⁴⁷ infrared spectroscopy cannot unequivocally identify the products after photoexcitation.⁴⁶ Therefore, an alternative, isomer-selective technique is needed.

2.2 PI-ReTOF-MS. Photoionization reflectron time-of-flight mass spectrometry (PI-ReTOF-MS) was exploited to identify the subliming molecules *isomer-specifically* during temperature programmed desorption (TPD) of the photoactivated ices from 5 to 300 K.⁴⁸⁻⁴⁹ This technique has the exceptional advantage to distinguish structural isomers based on their distinct adiabatic ionization energies (IEs) by systematically tuning the photon energy (PE) above and below the IE of the isomer of interest. As shown in Supporting Information (SI) Figure S2, The IEs of propynal (HCOCCCH , **7**), cyclopropanone (*c*- $\text{C}_3\text{H}_2\text{O}$, **3**), and propadienone (H_2CCCO , **8**) are 10.52–10.62 eV,⁵⁰ 9.22–9.32 eV,³³ and 9.04–9.14 eV,⁵¹ respectively, corrected for the Stark effect by the electric field of the ReTOF-MS optics.⁵² Therefore, two photon energies of 10.49 eV and 9.20 eV are required to detect cyclopropanone. Note that the ionization energies of the isotopically labelled (^{18}O , D) products differ by at most 0.01 eV from the non-substituted counterparts (**3**, **7**, **8**); this is within the experimental error limits.⁵² At 10.49 eV, only cyclopropanone (**3**), and propadienone (**8**) can be ionized; the IE of propynal (**7**) is higher. Lowering the PE to 9.20 eV would only allow – if present – ionization of propadienone (**8**). Therefore, through a comparison of the TPD profiles of the $\text{C}_3\text{D}_2^{18}\text{O}$ ions at $m/z = 58$ at distinct PEs of 10.49 eV and 9.20 eV, the formation of **3** versus **8** can be explored at distinct photoactivation wavelengths. Figure 2 compiles the PI-ReTOF-MS data of the molecules desorbing from the $\text{C}^{18}\text{O} - \text{C}_2\text{D}_2$ ices photoactivated at 222 nm, 249 nm, and 288 nm; Figure 3 displays the extracted TPD profiles of $m/z = 58$ ($\text{C}_3\text{D}_2^{18}\text{O}^+$) as a function of temperature at distinct photoactivation wavelengths and PEs (10.49 eV, 9.20 eV). Upon photoionization at 10.49 eV, a single sublimation event is evident from 150 K to 230 K for $m/z = 58$ of

the subliming ices photoactivated at 222 nm (Figure 3a). Recall that at PE = 10.49 eV, only **3** and **8** can be ionized. This sublimation event disappears when the PE is lowered to 9.20 eV (Figure 3b), which would only ionize **8** (Figure 3b). The blank experiment (Figure 3d) was conducted under identical conditions at PE = 10.49 eV, but without photoactivation of the ices. Therefore, we conclude that at 222 nm cyclopropenone ($c\text{-C}_3\text{D}_2^{18}\text{O}$, **3**) is formed. It should be stressed that the corresponding blank experiment (Figure 3d) shows no ion counts at all; hence ions at $m/z = 58$ are the result of the photoactivation of the ices at 222 nm, but they do not originate from ion-molecule reactions in the gas phase. The sublimation range of cyclopropenone ($c\text{-C}_3\text{D}_2^{18}\text{O}$, **3**) of 150–230 K agrees exceptionally well with the sublimation temperatures of cyclopropenone **3** prepared through the exposure of acetylene – carbon monoxide ices to energetic electrons at cryogenic temperatures (Figure 3a), while the propynal isomer **7** sublimes earlier from 128 – 155 K.⁵²⁻⁵³ Separate sets of experiments were conducted to photoactivate the $\text{C}^{18}\text{O} - \text{C}_2\text{D}_2$ ices at 222 nm over 0 (blank), 1, 2, 5, and 10 hours with the irradiation time plotted versus the integrated ion counts at $m/z = 58$ (Figure 4). This graph reveals a linear dependence of the cyclopropenone yield versus time, indicating that a single photon process is involved in the formation of $c\text{-C}_3\text{D}_2^{18}\text{O}$ with overall yields of $7.5 \pm 1.0 \times 10^{-19}$ ion counts per photon at 222 nm. Increasing the wavelength from 222 nm to 249 nm and photo-ionizing the subliming molecules at 10.49 eV also depicts a TPD profile from 150 – 230 K at $m/z = 58$ for cyclopropenone, albeit at lower count rates despite an increase in photon flux by a factor of three (Figure 3c). Compared to 222 nm, the ion yield is reduced to $9.7 \pm 1.3 \times 10^{-20}$ ion counts per photon at 249 nm. Increasing the wavelength further to 288 nm does not result in any ion counts above noise level despite an enhancement of the flux by a factor of four. Therefore, we find that, within our signal-to-noise limits, cyclopropenone is not formed upon photoactivation of $\text{C}^{18}\text{O} - \text{C}_2\text{D}_2$ ices at 288 nm. Based on the integrated ion counts and the photoionization cross sections,⁵² formation rates of cyclopropenone are determined to be $6.9 \pm 1.1 \times 10^{-2}$ and $8.9 \pm 1.4 \times 10^{-3}$ molecules per photon for 222 nm and 249 nm, respectively (Supporting Information).

2.3. Electronic Structure Calculations. The PI-ReTOF-MS experiments indicate that cyclopropenone is formed photochemically after $\text{C}^{18}\text{O} - \text{C}_2\text{D}_2$ ices are exposed to photons having a wavelength below 249 nm. We employ quantum chemical calculations to investigate the formation mechanism of $c\text{-C}_3\text{D}_2^{18}\text{O}$ starting from different low-lying excited states of C_2D_2 . The calculations are performed for a gas phase model of carbon monoxide and acetylene so that multi-

reference *ab initio* electronic structure methods can be used to determine ground and excited-state energetics of the cyclopropanone formation reaction. Details of the computations are included in Section 5.2 and in the Supporting Information. Briefly, the complete active space (CASSCF) approach⁵⁴ and second-order perturbation energy corrections (CASPT2//CASSCF)⁵⁵ with the ANO-L-VDZP basis set⁵⁶⁻⁵⁷ are employed for most calculations in this work. The formation mechanism of cyclopropanone was first investigated by performing relaxed potential energy scans along a decreasing distance between the carbon atom of carbon monoxide and the acetylenic carbon atom. On the ground state potential energy surface, calculations indicate an entrance barrier of 65 kJ mol⁻¹; i.e., the barrier for formation of *c*-C₃H₂O from the reactants CO(X¹Σ⁺) and C₂H₂(X¹Σ_g⁺). This is consistent with the fact that no *c*-C₃D₂¹⁸O was observed after heating carbon monoxide – acetylene ices from 5 K to 300 K in the “blank” experiment.

Next, the reaction between excited states of acetylene with ground state carbon monoxide are considered. For the experiments at 288 nm (4.31 eV) excitation, only the lowest triplet state T1 is energetically accessible (*cis*-C₂H₂ (a³B₂); *trans*-C₂H₂ (a³B_u)); at 249 nm (4.98 eV), the second excited triplet state of acetylene T2 can be accessed (*cis*-C₂H₂ (a³A₂); *trans*-C₂H₂ (a³A_u)); as for the 222 nm (5.58 eV) excitation, the third excited triplet state T3 ((*cis*-C₂H₂ (b³B₂); *trans*-C₂H₂ (b³B_u)) and the first excited singlet state S1 are energetically accessible ((*cis*-C₂H₂ (A¹A₂); *trans*-C₂H₂ (A¹A_u)). Both concerted and stepwise mechanisms for the formation of cyclopropanone are explored (Scheme 4). The distances between the carbon atom of carbon monoxide and the carbon atoms of acetylene are labeled as *d*_s for the concerted mechanism and *d*₁/*d*₂ for the stepwise mechanism. The concerted mechanism maintains C_{2v} symmetry during the reaction and the stepwise mechanism has a reduced C_s symmetry. While the stepwise mechanism may also proceed without maintaining planar symmetry, CASPT2//CASSCF calculations indicated that the reactions have similar energetics when computed with or without enforcing C_s symmetry. Since the electronic states are easier to identify when symmetrized, C_s symmetry is used for all stepwise mechanism calculations.

Although an excitation to either the *cis* or the *trans* isomer for each excited state is energetically possible, the *trans* isomers cannot react in a concerted mechanism that maintains C_s symmetry. Furthermore, CASPT2//CASSCF calculations without symmetry constraints rule out a stepwise mechanism that involves a *trans* isomer, since a triplet intermediate, where bond *d*₁ is formed, was found to be unstable in the singlet ground state and undergoes barrierless dissociation to carbon

monoxide and acetylene (Figure S3). Therefore, acetylene must be excited into a state in which the hydrogen atoms are oriented in a *cis* geometry. In the remainder of this section, using relaxed potential energy scans to simulate the concerted and stepwise mechanisms, we will discuss the reactivity of ground state carbon monoxide with excited states of *cis*-acetylene in its lowest three triplet (T1, T2, and T3) and the first singlet (S1) excited states.

2.3.1. Concerted Reaction Pathways. In all four excited low-lying *cis* states (T1, T2, T3, and S1), the excited state potential energy increases as carbon monoxide approaches perpendicularly to the acetylene carbon-carbon bond (Figure 5). The concerted mechanism is not energetically feasible for the T2, T3 and S1 states (a^3A_2 , b^3B_2 , A^1A_2) under our experimental conditions since the energy along the reaction coordinate rises above 5.59 eV (222 nm) before crossing with a higher excited state of the same symmetry. The only state from which a concerted mechanism is energetically accessible is the T1 state (a^3B_2) since it comes close to the ground state before rising in energy. However, the formation of cyclopropanone from the triplet excited state T1 (a^3B_2) requires intersystem crossing at a repulsive region of the T1 potential energy surface (PES). The spin-orbit coupling constant (SOCC) at the near-degeneracy at 2.2 Å ds is weak (1.79 cm^{-1} between the S1 and $m_s=\pm 1$ components of T1), only increasing to after leaving the near-degeneracy region of the PES (9.1 cm^{-1} at 1.8 Å). Furthermore, the sloped topology of the T1 and ground state crossing region indicates that a concerted reaction is unlikely, and that intersystem crossing to the ground state would lead back to the unreacted reactants.

The results of the potential energy scans can be rationalized upon inspection of the orbital symmetries of the reactants and the cyclopropanone product (Figure 6). The overall reaction involves a change of two electrons from A_1 to B_2 symmetry. While this could be achieved through an excited state with B_2 symmetry, one of the reactant A_1 orbitals becomes high in energy in the *c*- C_3H_2O product. A molecular orbital correlation would find that the reaction is effectively *symmetry-forbidden* when considering orbital reflection symmetry through the C_{2v} mirror planes.

2.3.2. Stepwise Reaction Pathways. The relaxed scans were repeated using C_s symmetry instead of C_{2v} such that only one distance, d_1 , is constrained (Scheme 4). The potential energies along the relaxed scans for T1, T2, T3, and S1 (now labeled as a^3A' , a^3A'' , b^3A' , and A^1A'' for C_s symmetry, and corresponding to a^3B_2 , a^3A_2 , b^3B_2 , and A^1A_2 for C_{2v} symmetry, respectively) are shown in Figure 7. Except for T3 (b^3A'), all states ultimately decrease in energy at small distances d_1 leading

to an excited-state minimum that is (near)-degenerate with the ground state. In all cases, the excited-state minima have a structure where the carbon monoxide carbon is attached to one of the acetylene carbon atoms. This is in contrast to the concerted (C_{2v}) mechanism where these same states become unstable at short distances.

Overall, the potential energy scans suggest that the triplet states T1 and T2 (a^3A' and a^3A'') and singlet S1 (A^1A'') can decay through intersystem crossing or non-adiabatic coupling, respectively, to a ground state structure that already has one OC–C₂H₂ bond formed. On the other hand, the triplet state T3 (b^3A') would have to undergo internal conversion to the lower a^3A' triplet state before reaching the ground state. The computed spin orbit coupling (SOC) between the T2 (a^3A'') and singlet ground state at the T2 optimized geometry is 11.35 cm⁻¹. In contrast, the SOC between T1 (a^3A') and S0 does not exceed 2 cm⁻¹ in the near-degeneracy region. While there is no barrier for OC–C₂H₂ bond formation on the T1 state, the T2 and S1 states appear to have a barrier due to an avoided crossing with a higher excited state of the same symmetry prior to arriving at the excited-state minimum. Such a crossing must be energetically accessible at the photon energies used. The relaxed scans indicate a barrier on the order of 0.5 eV for the T2 (a^3A'') state and 0.2 eV for the S1 (A^1A'') state, but those are upper limits due to symmetry and geometry constraints used in the calculations.

If the low-lying T1 state forms an intermediate structure where a carbon-carbon single bond is formed between the carbon monoxide and acetylene, why is no product observed after experimentally exciting with 288 nm photons? Recall that cyclopropanone is only detected upon excitation with 249 nm and 222 nm photons, which means that the product can only be formed from the T2 and S1 states corresponding to the a^3A_2 and A^1A_2 states under C_{2v} symmetry (Figure S4). To answer this question, we continue to map the reaction coordinate using a minimum energy path (MEP) calculation on the ground state PES starting from the T1 and T2 minimum-energy geometries (Figure 8). Due to the close resemblance of the S1 and T2 minimum energy geometries, the MEP from the S1 is not included in the figure since it yields a similar energy profile as T2. An important difference between T1 and T2 is the geometry of the minimum-energy structure (the triplet intermediate). For T2 (a^3A'') as well as S1 (A^1A''), the C=O bond is collinear with the OC–C₂H₂ bond (Figure 8 bottom). On the other hand, the T1 (a^3A') optimized structure has a carbon monoxide unit bonded at an OCC angle of 128°, parallel to the acetylene carbon-carbon bond (Figure 8 top). These different structures reflect distinct electronic characters of both states.

Running an MEP calculation starting from the T2 (a^3A'') and S1 (A^1A'') intermediate structures yields cyclopropenone ($c\text{-C}_3\text{H}_2\text{O}$, **3**) (Figure 8 bottom). However, running an MEP starting from the T1 (a^3A') minimum structure quickly converges to an unstable intermediate with a similar structure having slightly modified bond lengths and OCC angle 126° . A relaxed scan on the ground state PES along an increasing d_1 bond length indicates that the dissociation of carbon monoxide from that intermediate is barrierless at the CASPT2//CASSCF level of theory (Figure 8 top).

4. CONCLUSIONS

The joined experimental and computational investigation on the photolysis of acetylene – carbon monoxide ices provides persuasive testimony on *stepwise* cheletropic reaction pathways to cyclopropenone ($c\text{-C}_3\text{H}_2\text{O}$, **3**) via excited state dynamics involving the T2 (a^3A'') or S1 (A^1A'') states of acetylene. On the other hand, the reaction on the T1 (a^3A') state PES energetically favors an intermediate structure with a carbon-carbon single bond formed between carbon monoxide and acetylene that dissociates after relaxing to the ground state. The computational results agree well with the experiments, which reveal that cyclopropenone, as identified isomer-selectively via photoionization (PI) coupled with reflectron time-of-flight mass spectrometry (Re-TOF-MS), is only formed upon excitation of acetylene through a single-photon process with non-ionizing 222 nm and 249 nm photons, where S1 and T2 are energetically accessible, respectively, but not at 288 nm, where only T1 is energetically accessible. The overall reaction to yield $c\text{-C}_3\text{H}_2\text{O}$ from acetylene and carbon monoxide is endoergic by 57 kJ mol^{-1} , which can be compensated by the excess energy of the photons. The strong agreement between the experiments in low temperature ices and the outcome of excited state calculations represents a key benchmark of how computational and experimental photochemistry can jointly be exploited as a tool to gain molecular-level insight into excited-state reaction mechanisms and untangle exotic non-equilibrium chemical processes.

The photochemistry leading to the formation of $c\text{-C}_3\text{H}_2\text{O}$ is not only interesting from the fundamental viewpoints of physical organic chemistry (e.g., as unconventional non-equilibrium pathways to the formation of 2π - and 6π -aromatic ketones), but also reveals key insight into chemistry in more exotic environments such as ice-coated interstellar nanoparticles (interstellar grains) in cold molecular clouds and star-forming regions such as the Taurus Molecular Cloud (TMC-1) and Sagittarius-B2, (Sgr-B2), respectively.⁴⁷ Complex organic molecules (COMs) –

organics containing more than six atoms of carbon, hydrogen, nitrogen, and oxygen such as aldehydes and ketones - are ubiquitous in the interstellar medium (ISM), but their formation mechanisms have remained largely obscure. While gas phase astrochemical reaction networks replicate the abundance of hydrogen-deficient organics like linear cyanopolyynes,⁵⁸ pathways to complex organics are not understood. Since organic molecules constitute nearly 80 % of all detected interstellar molecules, the untangling of the fundamental processes leading to these COMs is crucial to unravel the basic mechanisms that initiate and drive low temperature organic chemistry in space. Therefore, by exploiting cyclopropenone as a benchmark, laboratory astrophysics experiments coupled with theoretical calculations, as presented here, expose a hitherto overlooked reaction class leading to COMs via photon-triggered (non-adiabatic) excited state reactions. These processes may occur at ultralow temperatures (10 K) within interstellar ices, rather than through traditional radical-radical pathways on grain surfaces in the warm-up phase of the ices as hypothesized for the last decades, changing our perception on the formation of complex organics in the interstellar medium and eventually leading to the molecular precursors of biorelevant molecules as planets form in their interstellar nurseries.

5. EXPERIMENTAL AND COMPUTATIONAL

5.1 Experimental Section. The experiments were conducted in a contamination-free ultrahigh vacuum chamber pumped to a few 10^{-11} Torr.⁴⁵ A 1 cm^2 silver mirror, which served as the substrate, was interfaced to a cold finger was cooled to $5.0 \pm 0.1 \text{ K}$ by a closed-cycle helium compressor (Sumitomo Heavy Industries, RDK-415E). The $\text{C}^{18}\text{O} - \text{C}_2\text{D}_2$ ices were prepared by depositing gas mixtures of isotopically labeled ^{18}O -carbon monoxide (C^{18}O , 95% ^{18}O , Sigma Aldrich) and D_2 -acetylene (C_2D_2 , 99% D, CDN isotope) via a glass capillary array at a background pressure of 5×10^{-8} Torr until ice thickness was achieved.⁵⁹ Since C_4H_6 isomers (54 amu) were detected in irradiated acetylene ices⁶⁰ and may overlap with cyclopropenone ($\text{C}_3\text{H}_2\text{O}$; 54 amu) formed by $\text{CO} - \text{C}_2\text{H}_2$ ices, isotopically labeled $\text{C}^{18}\text{O} - \text{C}_2\text{D}_2$ ices were used to be able to distinguish D_2 - ^{18}O cyclopropenone ($\text{C}_3\text{D}_2^{18}\text{O}$; 58 amu) and its isomers. An ice thickness of $960 \pm 30 \text{ nm}$ was determined *in situ* using an estimated index of refraction of the mixed $\text{C}^{18}\text{O} - \text{C}_2\text{D}_2$ ice of 1.32 ± 0.03 . FTIR (Nicolet 6700) was used to monitor the $\text{C}^{18}\text{O} - \text{C}_2\text{D}_2$ ice mixture at 5 K both online and *in situ* before, during, and after the irradiation phase of the experiment in the range of 500 to 6000 cm^{-1} with a resolution of 4 cm^{-1} in intervals of approximately 2 minutes. After the deposition, the $\text{C}^{18}\text{O} - \text{C}_2\text{D}_2$ ice was photoexcited over a $1.0 \pm 0.1 \text{ cm}^2$ area at an angle of incidence of 0° with

respect to the surface normal of the substrate at wavelengths of 222 nm, 249 nm, and 288 nm with average laser powers of 10 mW. After the irradiation, the ice was sublimed via TPD as the substrate was heated to 300 K at 0.5 K min⁻¹. The analysis of any newly formed products in the gas phase after irradiation was achieved by the PI-ReTOF-MS.⁴⁶ More experimental details are included in the Supporting Information.

5.2 Computational Section. To capture the multi-reference character of electronically excited acetylene during its reaction with CO, geometry optimizations were carried out with the complete active space self-consistent field (CASSCF) method,⁵⁴ while single-point energy corrections are performed with the complete active space second-order perturbation theory (CASPT2) level of theory.⁵⁵ An active space of 10 electrons in 9 orbitals is employed for all calculations, shown in Figure 6. In addition to symmetry constraints, relaxed scans were carried out by constraining the distance between the CO carbon atom and one of the C₂H₂ carbon atoms (d_s or d_1/d_2 in Scheme 4) at different values. Minimum energy path (MEP) calculations employed a step size of 0.05 Bohr and a maximum step size of 0.02 Bohr. All CASPT2 calculations employed an imaginary level shift of 0.2.⁶¹ No Ionization Potential-Electron Affinity (IPEA) shift was used.⁶²⁻⁶³ CASSCF geometry optimizations and CASPT2 energy calculations were carried out using OpenMolcas version 21.06.⁶⁴ Natural transition orbitals (NTOs, Figure S3)⁶⁵⁻⁶⁶ for acetylene in the linear and bent (*cis* and *trans*) geometries were computed with time-dependent density functional theory using the B3LYP functional and 6-311++G** basis set. NTO calculations were performed with Q-Chem and visualized in IQmol.⁶⁷ Additional details of the computational methods are given in the Supporting Information.

Supporting Information

Experimental details, quantum chemical calculations, Table S1, Figures S1–S4, Cartesian coordinates of key structures, and absolute energies.

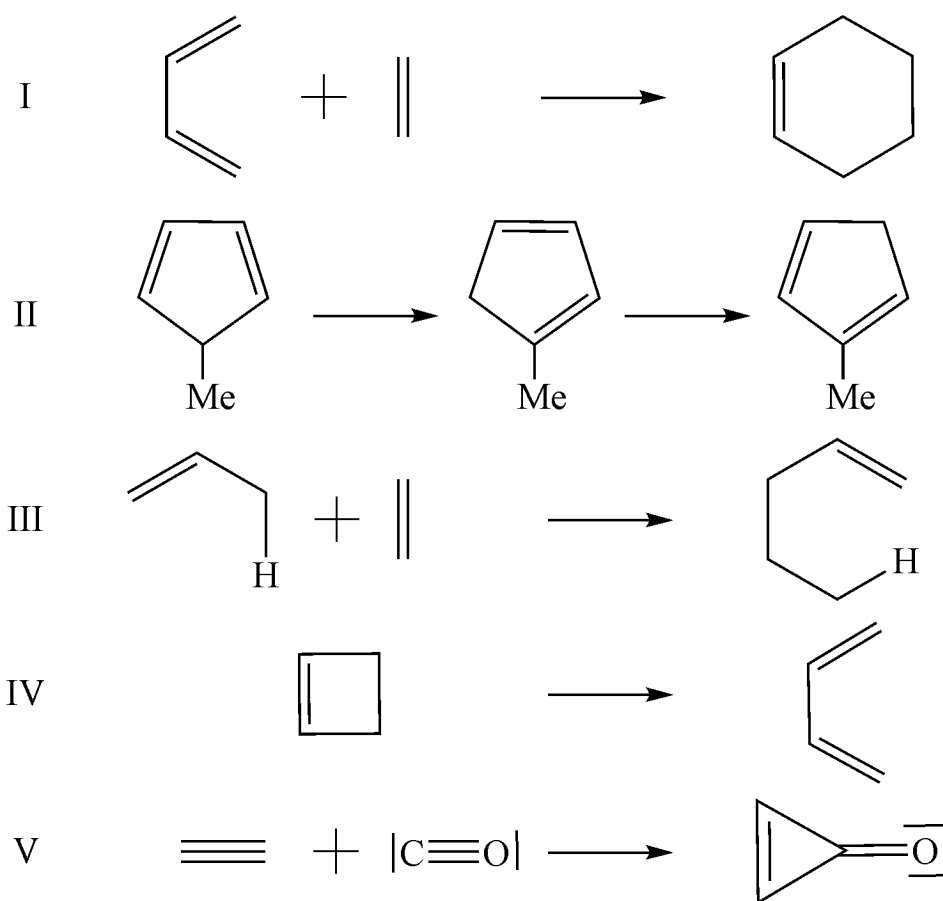
Author Contributions

The manuscript was written through the contributions of all authors. All authors have given approval to the final version of the manuscript.

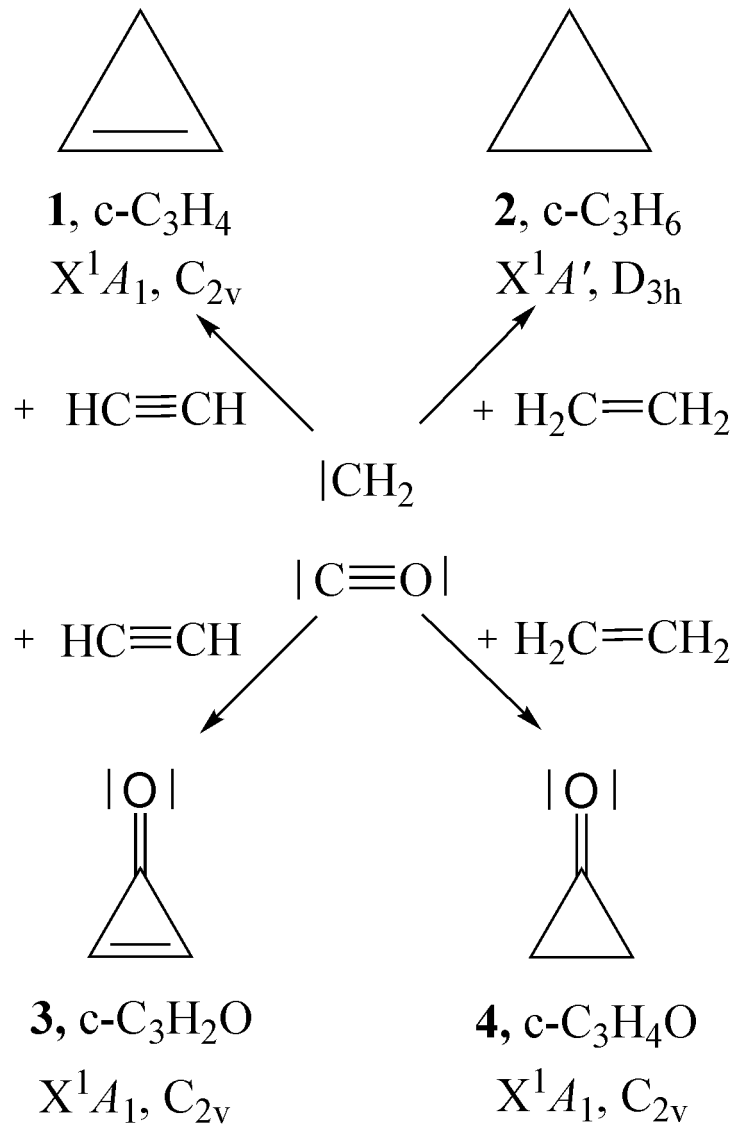
Notes

The authors declare no competing financial interest.

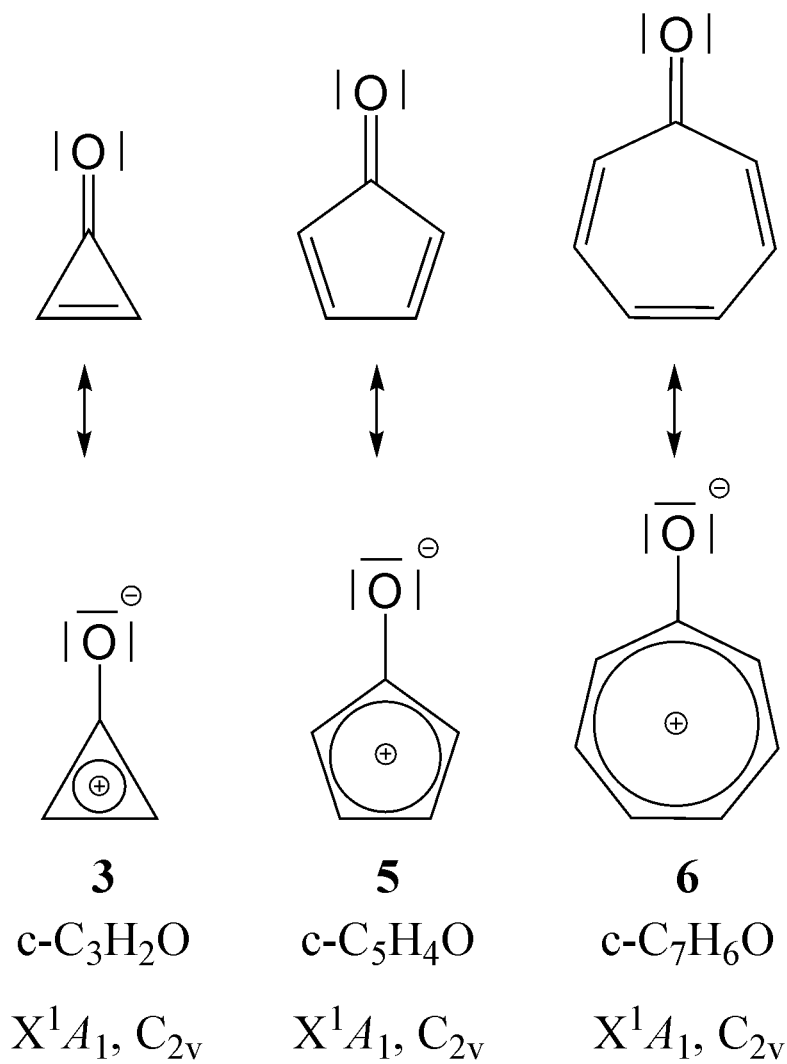
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Scheme 1. Examples of pericyclic reactions: cycloaddition (I), sigmatropic shift (II), group transfer reaction (III), electrocyclic reaction (IV), and cheletropic reaction (V).



Scheme 2. Schematic representation of cheletropic reactions of singlet carbene (CH₂) with acetylene (C₂H₂) and ethylene (C₂H₄) leading to cyclopropene (C₃H₄, **1**) and cyclopropane (C₃H₆, **2**) and of carbon monoxide (CO) with acetylene (C₂H₂) and ethylene (C₂H₄) forming cyclopropenone (c-C₃H₂O, **3**) and cyclopropanone (c-C₃H₄O, **4**), respectively.



Scheme 3. Electronic wave functions and point groups of cyclopropenone (**3**), cyclopentadienone (**5**), and cycloheptatrienone (**6**) along with their resonance structures.

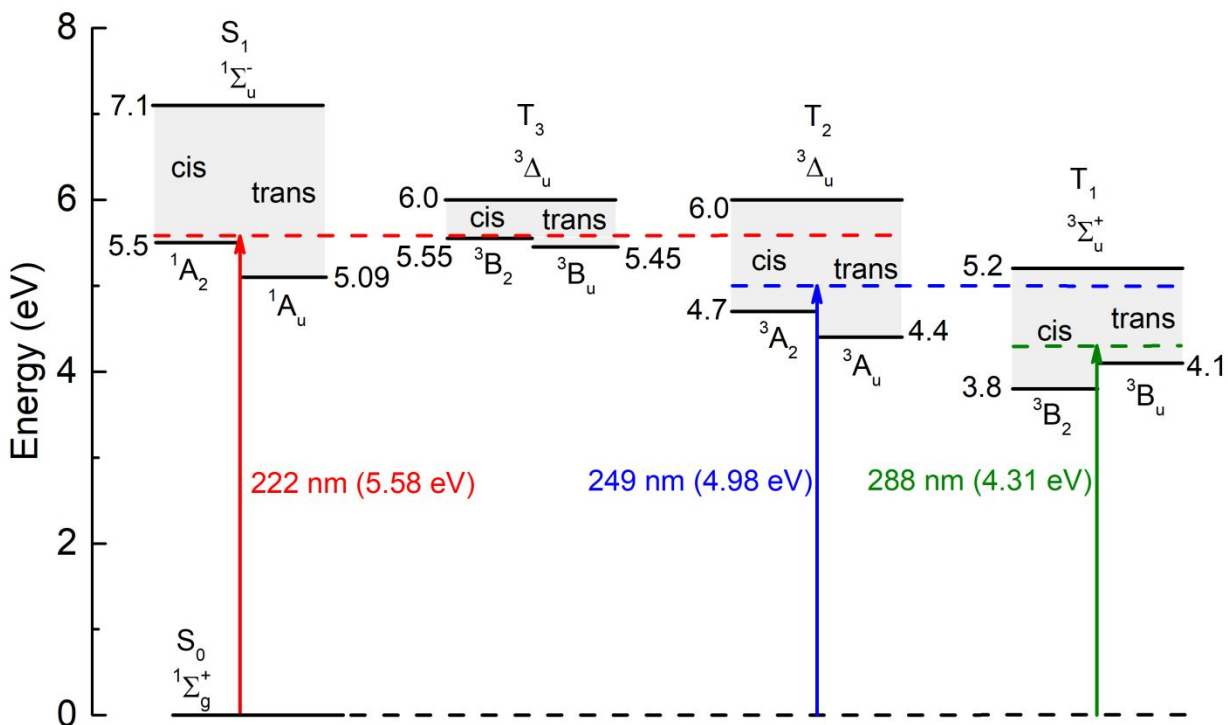


Figure 1. Energy level diagram of the four lowest energy excited states of acetylene (T_1 , T_2 , T_3 , and S_1).⁵² The upper and lower energy limits of each state correspond to the linear symmetry (vertical excitation) and the adiabatic transitions to the *cis* and *trans* geometries, respectively. Arrows and dashed lines represent the excitation energies used in the experiments.

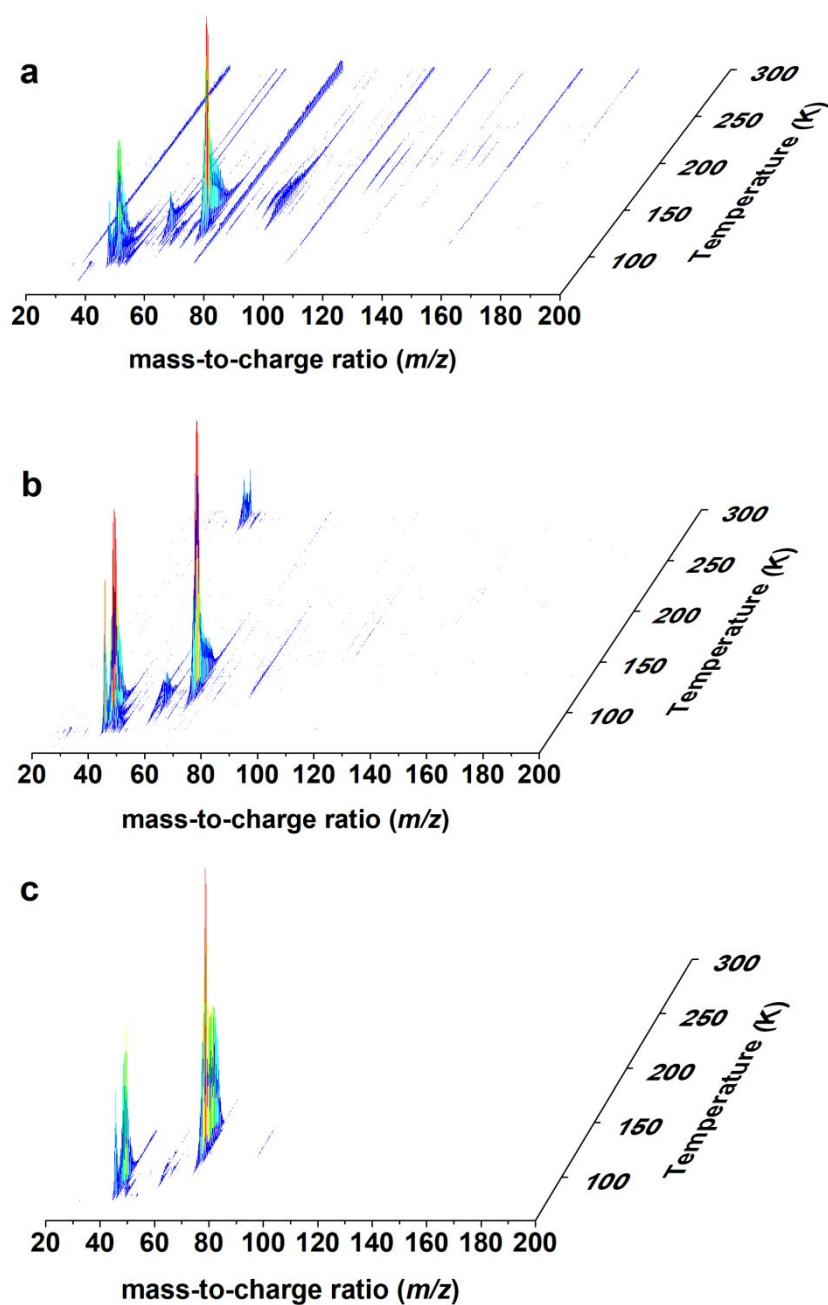


Figure 2. PI-ReTOF-MS data displaying ion counts as a function of the temperature of the subliming molecules recorded at 10.49 eV for ^{18}O -carbon monoxide – D₂-acetylene ices ($\text{C}^{18}\text{O} - \text{C}_2\text{D}_2$) photolyzed at 222 nm (a), 249 nm (b), and 288 nm (c).

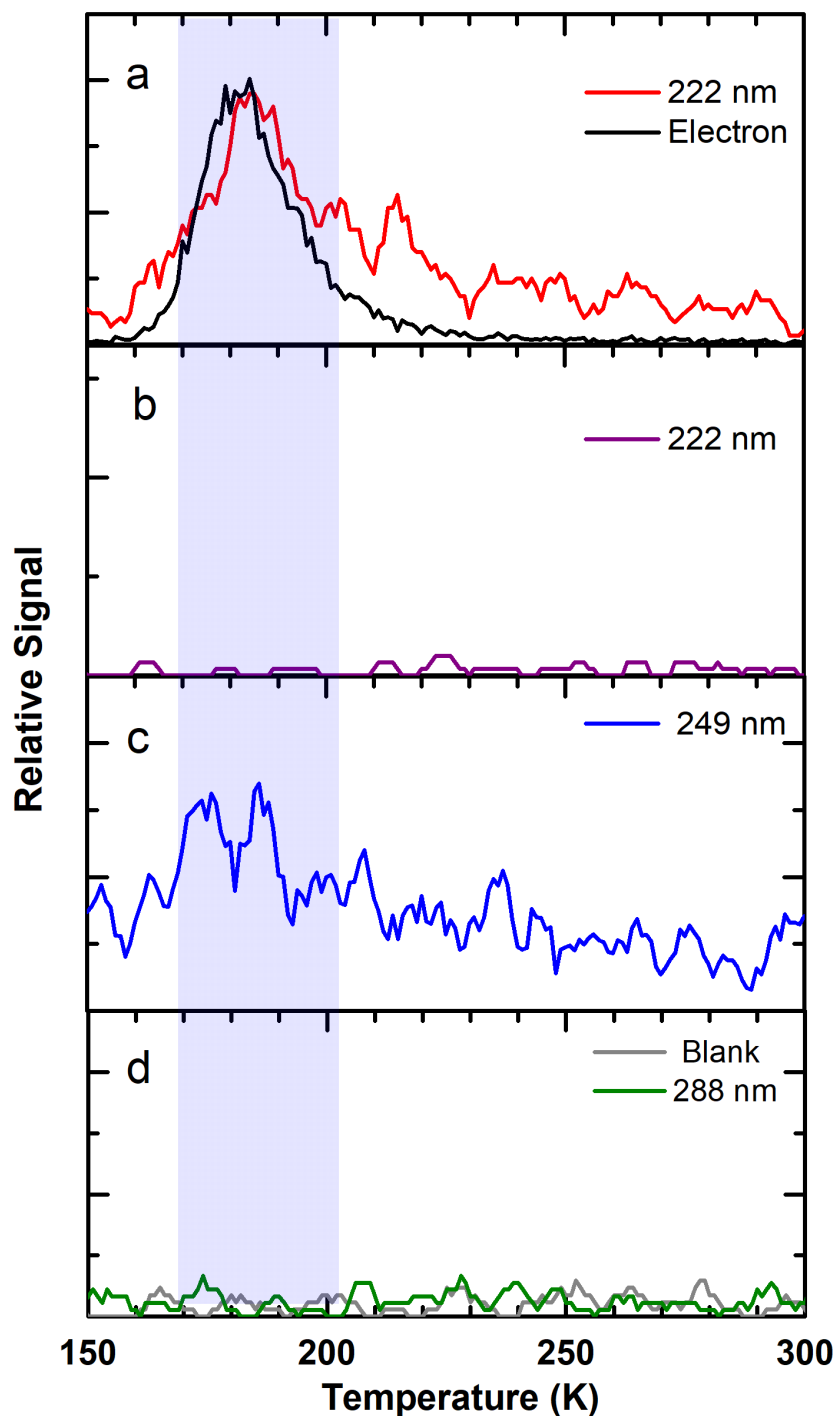


Figure 3. TPD profiles for $m/z = 58$ ($\text{C}_3\text{D}_2^{18}\text{O}^+$) of photolyzed ices recorded at photoionization energies of 10.49 eV (a, c, d) and 9.20 eV (b). The $\text{C}^{18}\text{O} - \text{C}_2\text{D}_2$ ices were photolyzed by 222 nm (purple and red) for 10 hours, by 249 nm (blue) for 30 hours, and by 288 nm (green) for 42 hours. The blank experiment (grey) was carried out under identical conditions, but without photolysis of the ices.

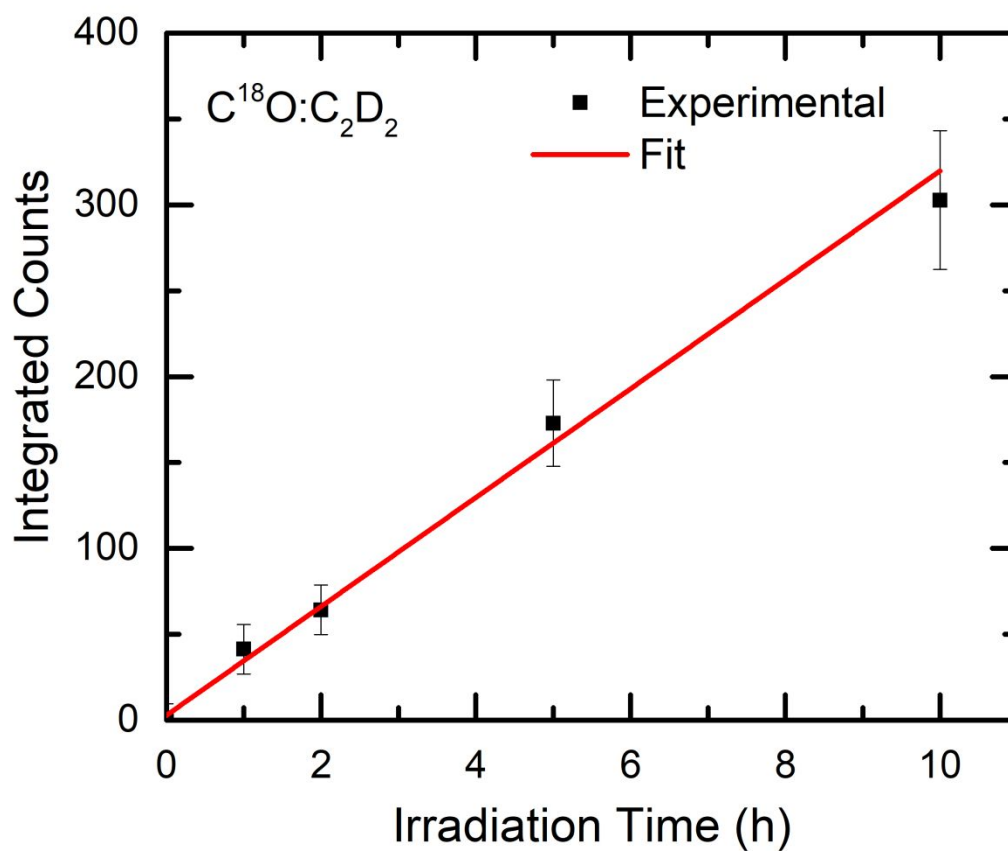
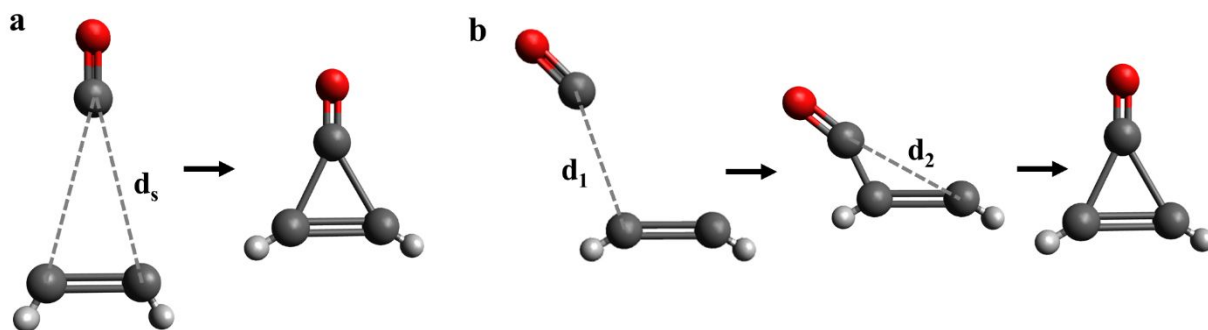


Figure 4. Linear relationship of integrated ion counts of m/z 58 collected during the TPD phase of the $C^{18}O - C_2D_2$ ices photolyzed at 222 nm versus the photolysis time.



Scheme 4. The proposed concerted (a, C_{2v} symmetry) and stepwise (b, C_s symmetry) mechanisms for the formation of *c*-C₃H₂O.

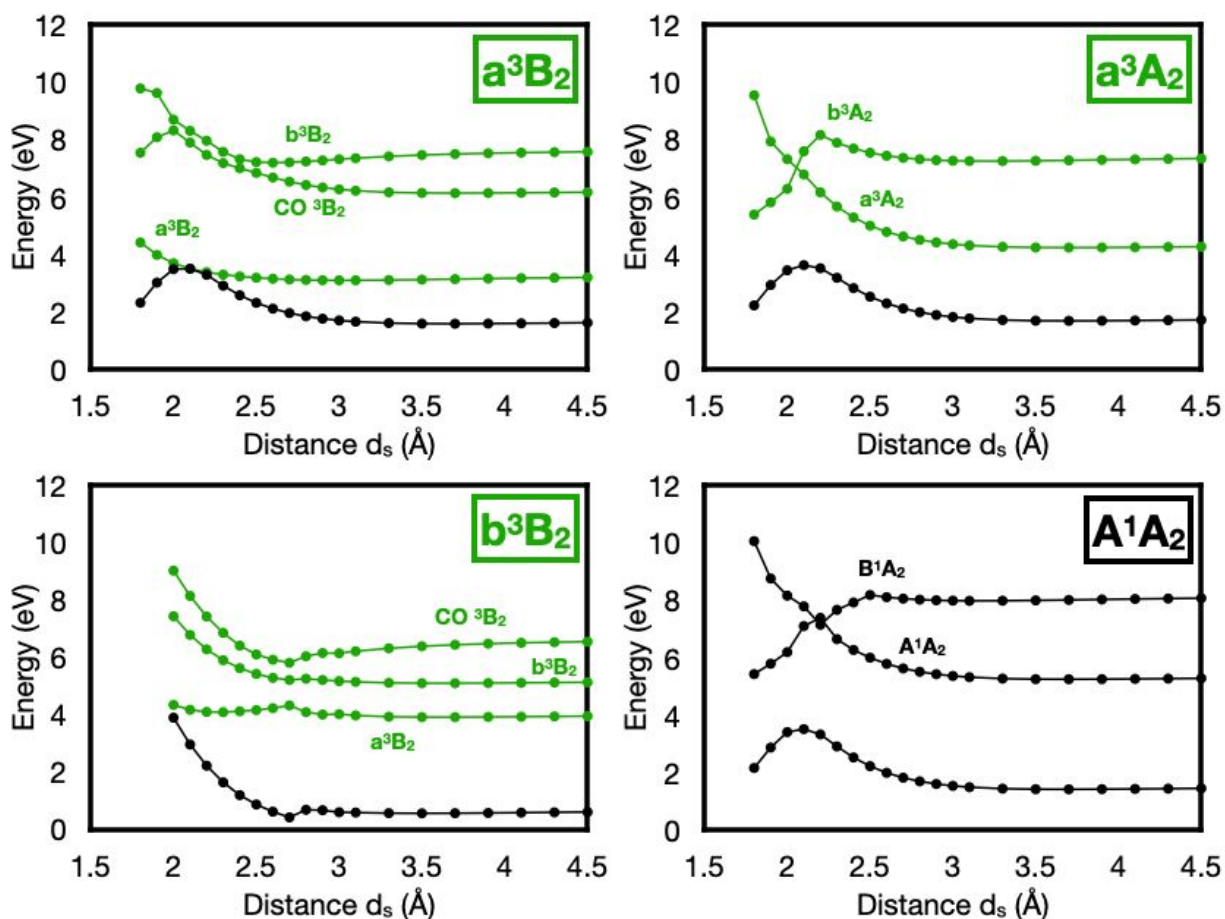


Figure 5. Potential energy scans along d_s (Scheme 4) for the four lowest adiabatic excited states of acetylene (T1, T2, T3, and S1, corresponding to a^3B_2 , a^3A_2 , b^3B_2 , and A^1A_2). These calculations were carried out enforcing C_{2v} symmetry. The state for which gradients are used for the relaxed scan is indicated on the top right of each panel. Triplet states are indicated in green, while singlet states are indicated in black. The lowest singlet excited state in each panel is the ground state. Energies are reported relative to the ground state equilibrium geometry of acetylene and carbon monoxide at 4.5 Å separation, i.e., with linear acetylene.

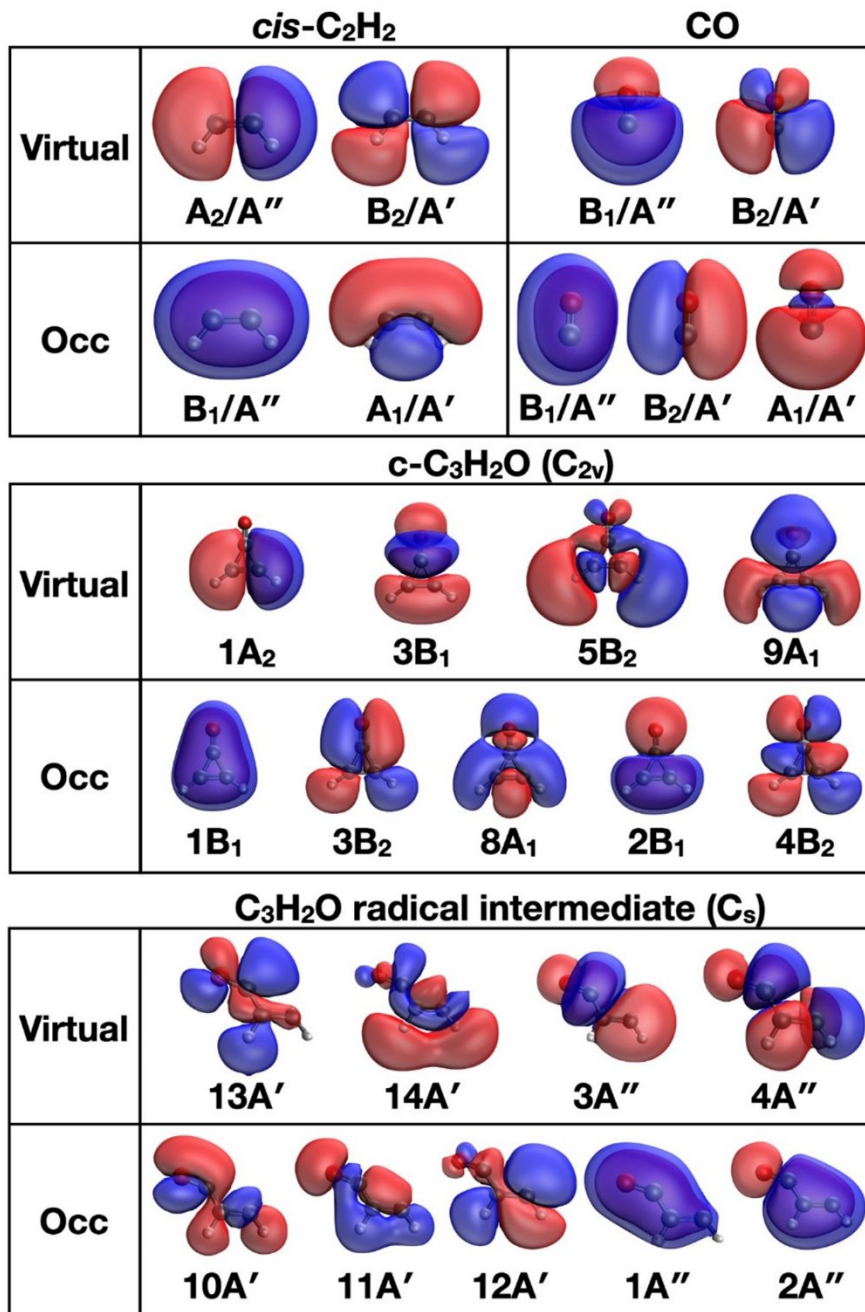


Figure 6. Nine molecular orbitals used in the active space for CASPT2 and CASSCF calculations, shown for the reactants (top), product (middle), and open-shell intermediate (bottom). Specifically, the orbitals are shown for an excited-state optimized structure of acetylene in the *cis* stereoisomer, for carbon monoxide, and *c*-C₃H₂O (**3**) in their ground-state optimized structures, and a C₃H₂O radical optimized using the triplet state of the A' irreducible representation (*a*³A'). The molecular orbital labels use the C_{2v} irreducible representation for *c*-C₃H₂O (**3**), the C_s irreducible representation for the intermediate, and both C_s and C_{2v} irreducible representation for the reactants. Occ is an abbreviation for “occupied.”

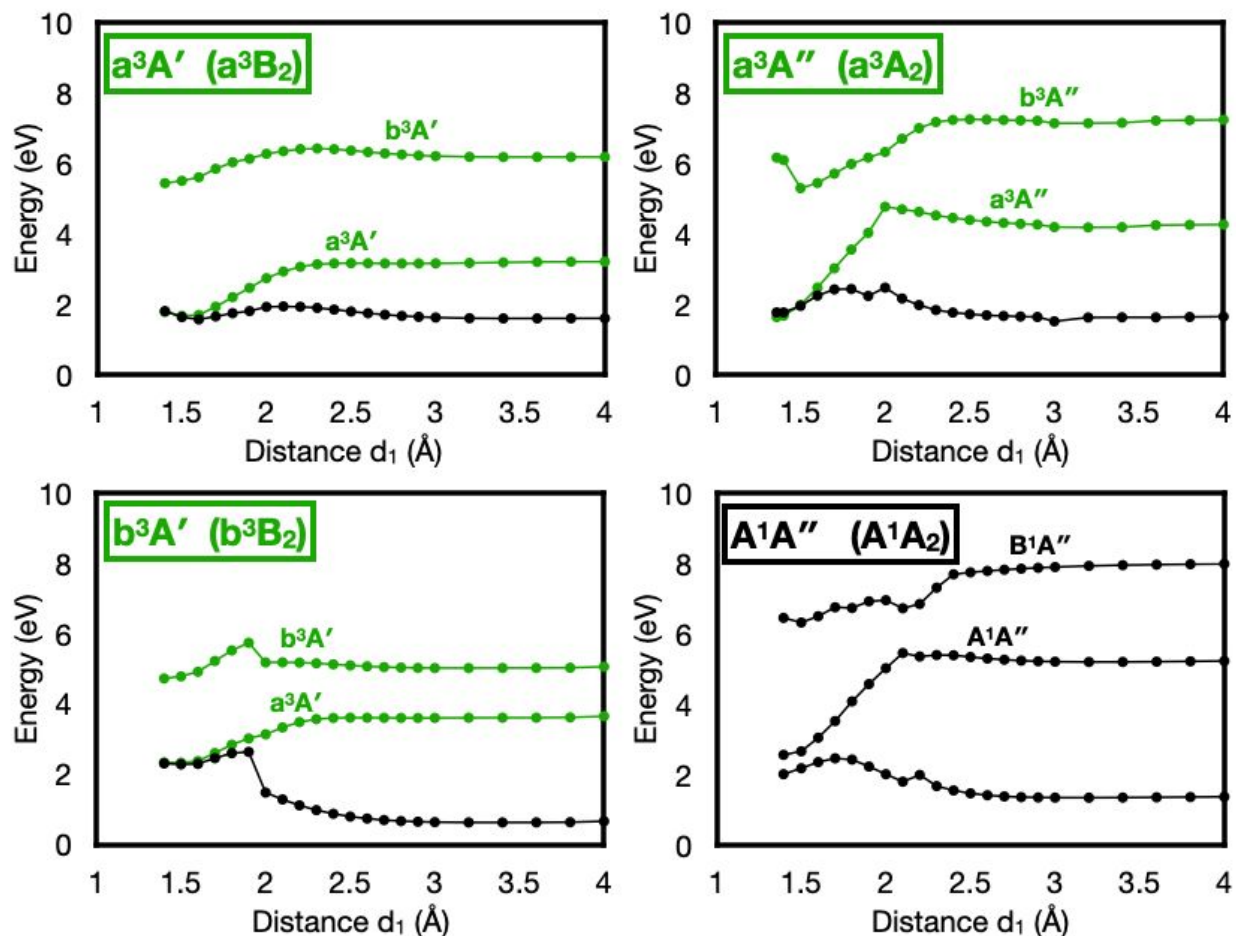


Figure 7. Potential energy scans along d_1 (Scheme 4) for the four lowest adiabatic excited states of acetylene (T1, T2, T3, and S1, corresponding to a^3A' , a^3A'' , b^3A' , and A^1A''). These calculations were carried out enforcing C_s symmetry. The state for which gradients are used for the relaxed scan is indicated on the top right of each panel; both C_s and C_{2v} labels are used to help comparison with earlier Figures. Triplet states are indicated in green, while singlet states are indicated in black. The lowest singlet excited state in each panel is the ground state. Energies are reported relative to the ground state equilibrium geometry of acetylene and carbon monoxide at 4.5 Å separation.

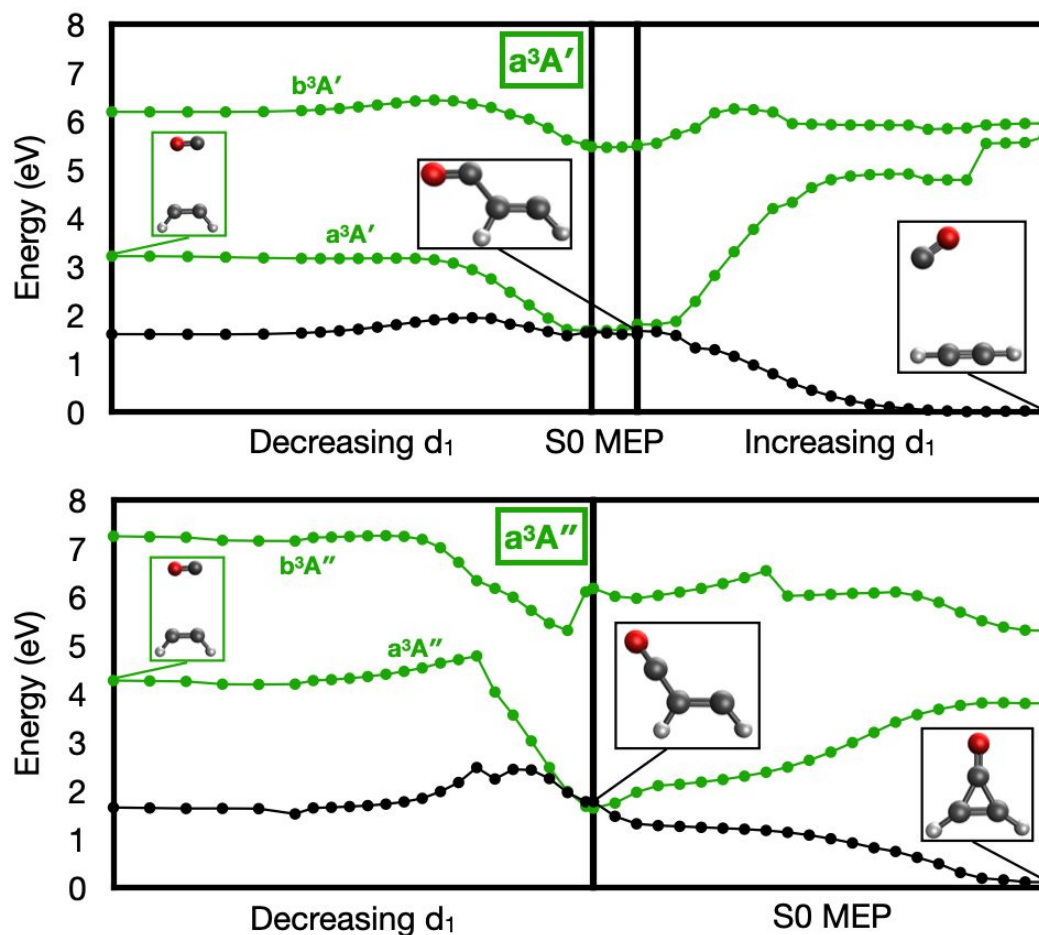


Figure 8. Top: A combination of three calculations for the T1 state (a^3A'). The left plot is adapted from Figure 5 and inverted to show decreasing rather than increasing d_1 . The central section is a minimum energy path (MEP) calculation on the ground state potential energy surface (PES), starting from the T1 (a^3A') triplet state minimum geometry. The plot on the right is a relaxed potential energy scan on the ground state PES while increasing the distance d_1 . **Bottom:** A combination of two calculations for the T2 (a^3A'') state. The left plot is adapted from Figure 7 but inverted to show decreasing rather than increasing d_1 . The plot on the right is a MEP calculation on the ground state PES, starting from the triplet state T2 (a^3A'') minimum geometry. All calculations were carried out enforcing C_s symmetry. Triplet states are indicated in green, while singlet states are indicated in black. The lowest singlet excited state in each panel is the ground state. Energies are reported relative to the ground state equilibrium geometry of acetylene and carbon monoxide at 4.5 Å separation, i.e., with a linear acetylene). Select structures for specific points are shown in boxes.

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