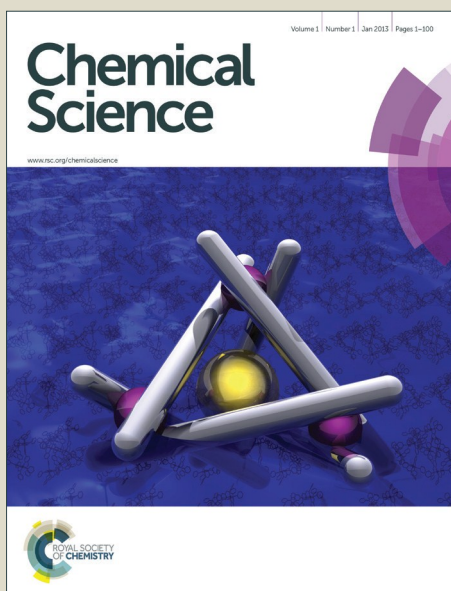


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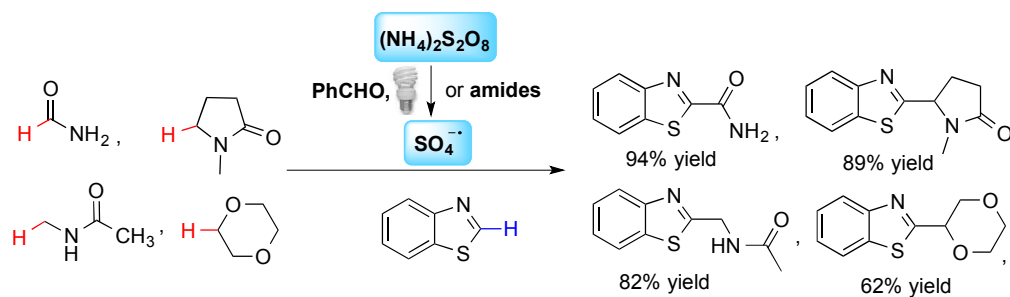
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Graphic Abstract

A benzaldehyde-mediated photoredox reaction for α -heteroarylation of amides (α to nitrogen) and ethers through cross-dehydrogenative coupling (CDC).





Journal Name

ARTICLE

Direct α -Heteroarylation of Amides (α to Nitrogen) and Ethers through a Benzaldehyde-Mediated Photoredox Reaction

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A photoredox reaction for cross-dehydrogenative coupling (CDC) was developed to C_{α} -arylate amides (α to nitrogen) and ethers using a variety of five- and six-membered electron-deficient heteroarenes. A unique decomposition mechanism of ammonium persulfate enhanced by photoexcited benzaldehydes was disclosed. This benzaldehyde-mediated photoredox reaction proceeded smoothly with household 23 W CFL bulbs as the energy source under metal-free conditions, allowing the construction of new C_{sp^2} - C_{sp^2} and C_{sp^3} - C_{sp^2} bonds and access to important pharmacophores of broad utility using commercially available reagents.

Introduction

Coupling reactions that connect nitrogen-containing heteroaromatic moieties to other molecular structures are a mainstay transformation in the synthesis of medicinal agents. The C–H bonds α to nitrogen or oxygen atoms in amines, amides (α to nitrogen), and ethers can be activated through a single electron transfer (SET) process to generate corresponding radicals.¹ This strategy enables a direct conversion of feedstock amines, amides, ethers, etc. into α -heteroarylated products that become important and broadly employed pharmacophores for drug discovery (see Fig. S1†).² However, the generation of such radicals and the subsequent coupling reactions require harsh conditions, such as high temperature, stoichiometric amounts of oxidants, and the use of transition-metal catalysts.¹

Recently, metal-based, visible light-driven photoredox catalysis has been identified as a suitable approach for the generation of α -carbon radicals of amines and ethers.³ This approach facilitates the coupling reactions of the generated radicals with aromatic heterocycles under mild conditions. One of the most fascinating aspects of this visible light-based photochemistry is that the reaction can be conducted with a small quantity of metal-based photocatalyst and a household compact fluorescent light (CFL) bulb, which is easier to handle than UV lamps. Further, complex organic molecules are more stable towards photodecomposition when they are irradiated with lower energy wavelengths.⁴ Nevertheless, the current methodologies require pre-functionalized heteroarenes (with CN and Cl) that consequently produce stoichiometric

amounts of undesired waste (Fig. 1a), and the coupling partners are limited to amines and ethers.³ The reactions for electron-poor amide substrates remain unexplored. Thus, the development of new photoredox cross-dehydrogenative-coupling (CDC) reactions, which can couple medicinally important heterocycles with a broader scope of coupling partners and selectively form C–C bonds between two different inert C–H bonds without any pre-functionalization, is highly desirable.⁵

Notably, MacMillan and co-workers recently reported a CDC-based reaction for electron-deficient heteroarenes

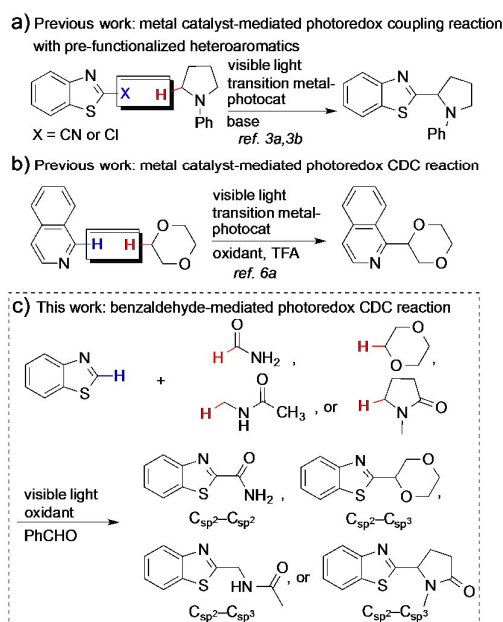


Fig. 1 Photoredox-based coupling reactions.

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† Electronic Supplementary Information (ESI) available: Fig. S1–S4, Scheme S1–S3, experimental procedures and characterization for all new compounds. See DOI: 10.1039/x0xx00000x

through a combination of iridium- or ruthenium-containing photoredox catalysis with a Minisci-type reaction of α -oxy radicals (Fig. 1b).^{6a} A similar strategy was also applied in the synthesis of a JAK2 inhibitor by Stephenson *et al.*^{6b} However, these reported reactions only coupled six-membered nitrogen-containing heteroaromatics with ethers or electron-rich amines. There is an urgent need to develop an efficient photochemically induced CDC reaction for a broader range of heteroaromatics with diverse amines, amides, and ethers. It would also be advantageous that the new reaction avoids the use of expensive transition metal catalysts or toxic/explosive reagents. Herein, we report a photoredox CDC reaction that enables the selective α -heteroarylation of amides (α to nitrogen) and ethers through C–H activation (Fig. 1c). This protocol features a distinct decomposition mechanism of APS promoted by photoexcited benzaldehydes, the metal-free reaction conditions using an environment-friendly oxidant radical initiator and 23 W CFL bulbs, and the coupling of amides and ethers with five- and six-membered electron deficient heteroarenes.

Results and discussion

(NH₄)₂S₂O₈ (ammonium persulfate, APS) is widely used as a cheap and environmentally benign radical initiator in organic synthesis and the polymer industry.⁷ In combination with a metal-based photocatalyst, such as Ru(bpy)₃Cl₂, visible light can drive the decomposition of APS to produce sulfate radicals, which are widely involved in the initiation of diverse radicals through the hydrogen atom transfer (HAT) process.^{6,8} This strategy has been successfully implemented in the CDC reactions of ethers with heteroaromatics.⁶ We questioned whether the α -amino radical could also be initiated via this photochemically induced process following by a Minisci-type reaction with electron-deficient heteroarenes to create a new CDC reaction.

Formamide has a relatively low C–H bond strength (BDE = 94 kcal/mol)⁹ that tends to promote a homolytic cleavage to generate a C_{sp2}-based nucleophilic α -amino radical. Moreover, a sunlight-induced coupling reaction of formamide with benzene-fused pyridines was successfully implemented with polycrystalline TiO₂ as the catalyst. This reaction exhibits low reaction efficiency and limited substrate scopes.¹⁰ Our examination of the proposed photochemically induced CDC reaction with heteroaromatics began with formamide and benzothiazole given the utility of benzothiazoles in biomedicine.¹¹ The initial reaction employing 2 mol% of Ru(bpy)₃Cl₂ provided the carbamoylated benzothiazole, however, with a low yield (entry 1 in Table 1, 6% yield). Surprisingly, the product was obtained in a moderate yield when the photocatalyst was excluded (entry 2, 59% yield). Lowering the amount of APS resulted in a decreased yield (entry 5, 25% yield). The careful exclusion of light had little effect on the reaction (entry 4, 55% yield), while decreasing the temperature led to lower reaction efficiency (entry 3, 15% yield). This result may be attributed to the temperature-dependent decomposition of APS enhanced by formamide,

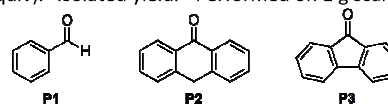
which has been described in the APS-initiated polymerization of acrylamide.¹²

Interestingly, even 0.5 equivalent of benzaldehyde can promote the reaction (entry 6 in Table 1, 90% yield). In the absence of CFL illumination, benzaldehyde did not improve the yield (entry 8, 3% yield). In addition, no product was detected when the reaction was performed with benzaldehyde but not APS (entry 9). These results indicated that the photoexcited benzaldehyde might be involved in this photoredox process to promote the decomposition of APS and the generation of the carbamoyl radical. Benzophenone, which is reported to mediate the SET process upon the illumination of CFL bulbs, can also improve the yield (entry 10, 84% yield). On the other hand, 9-fluorenone that displays a stronger absorption in the visible light region (Fig. S2) completely suppressed the enhanced reactivity (entry 11, 34% yield, *versus* entries 2 and 7 in Table 1).¹³ Therefore, 1.0 equivalent of benzaldehyde was selected for further optimization. Lowering the amount of formamide to 50 equivalents resulted in a decreased yield (entry 12, 46% yield). The reaction was inhibited in the presence of air (O₂) and TEMPO (entries 13 and 14), again suggesting a radical mechanism. Further, *p*-toluenesulfonic acid monohydrate (1 equiv) can improve the yield (entry 15, 99% yield) through protonation and activation of benzothiazole. In the presence of *p*-toluenesulfonic acid,

Table 1 Preliminary studies towards the carbamoylation of benzothiazole^a

Entry	(NH ₄) ₂ S ₂ O ₈	Photosensitizer	TsOH	Yield ^b
1 ^c	3 equiv.	–	–	6%
2	3 equiv.	–	–	59%
3 ^{d,e}	3 equiv.	–	–	15%
4 ^e	3 equiv.	–	–	55%
5	1 equiv.	–	–	25%
6	3 equiv.	P1 (0.5 equiv.)	–	90%
7	3 equiv.	P1 (1.0 equiv.)	–	95%
8 ^{d,e}	3 equiv.	P1 (1.0 equiv.)	–	3%
9	–	P1 (1.0 equiv.)	–	n.r. ^f
10	3 equiv.	P2 (1.0 equiv.)	–	84%
11	3 equiv.	P3 (1.0 equiv.)	–	34%
12 ^g	3 equiv.	P1 (1.0 equiv.)	–	46%
13 ^h	3 equiv.	P1 (1.0 equiv.)	–	13%
14 ⁱ	3 equiv.	P1 (1.0 equiv.)	–	n.r. ^f
15	3 equiv.	P1 (1.0 equiv.)	1.0 equiv.	99% (94% ^j)
16	1 equiv.	P1 (1.0 equiv.)	1.0 equiv.	70%
17	3 equiv.	–	1.0 equiv.	86%
18 ^k	3 equiv.	P1 (1.0 equiv.)	1.0 equiv.	85% ^j

^a General procedures A–C for the experimental details unless otherwise noted. ^b Yields were determined by the ¹H NMR spectroscopy using CH₂Br₂ as an internal standard. ^c Performed with Ru(bpy)₃Cl₂ (2 mol%). ^d Performed at 23 °C. ^e Performed in darkness. ^f No reaction. ^g Performed with HCONH₂ (50 equiv.). ^h Performed under air. ⁱ Performed with TEMPO (3 equiv.). ^j Isolated yield. ^k Performed on 1 g scale.

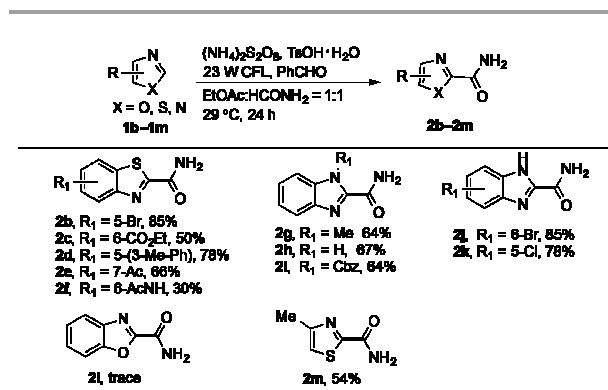


lowering the amount of APS (entry 16) or the exclusion of benzaldehyde (entry 17) led to decreased yields. Hence, the optimized reaction conditions are shown in entry 15. With these conditions in hand, the reaction was successfully implemented on a gram scale (entry 18, 85% isolated yield).

The synthetic potential of this new photoredox reaction was then evaluated with different heteroarene compounds. For benzothiazole substrates with electron-withdrawing or weak electron-donating groups, the reactions proceeded well (**2b–2e** in Table 2, 50–85% yields) with excellent selectivity at the C2 position. Bromide, ester, and acyl groups were well tolerated, providing handles for functionalization. However, the introduction of a stronger electron-donating group, such as the acetamido group, to benzothiazole resulted in diminished reactivity towards nucleophilic addition of the carbamoyl radical, lowering the yield (**2f**, 30% yield). These results further indicated the nucleophilic nature of the reaction. Benzimidazoles also functioned as suitable substrates (**2g–2k**, 64–85% yields). Benzoxazole was not as well tolerated for this protocol (**2l**, trace) with the reaction resulting in a complicated mixture perhaps owing to its instability to strong oxidants.¹⁴ Finally, the monocyclic thiazole was as well amendable to this coupling strategy (**2m**, 54% yield).

Having successfully demonstrated the carbamoylation of a variety of azole heterocycles, we anticipate that electron-

Table 2. Benzaldehyde-mediated photoredox CDC reactions of azoles and benzazoles with formamide^a



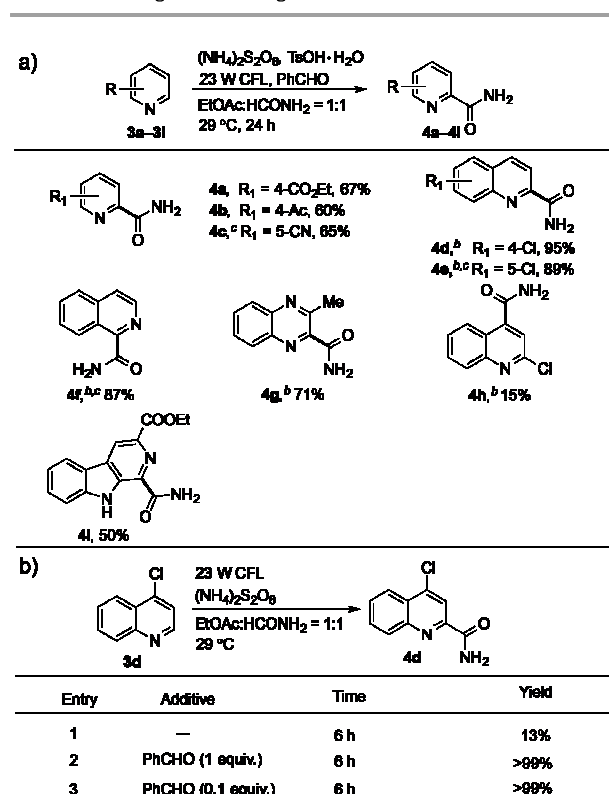
^a See general procedure A for the experimental details unless otherwise noted. Isolated yield was reported for each reaction.

deficient six-membered nitrogen-containing heteroaromatics can also be engaged with this new protocol (Table 3). The pyridine substrates underwent coupling with moderate yields (**4a–4c** in Table 3a, 60–67% yields). Pyridine or pyrazine fused with a benzene ring displayed increased reaction efficiency and produced excellent yields even without the acidic additive (**4d–4g**, 71–95% yields). For pyridine-containing substrates, *ortho* substitution is more favored than *para* substitution (**4c**, C-6:C-2:C-4 = 16.5:5.5:1; and **4e**, C-2:C-4 = 3:1). The bis-addition products proved to be minimal. The H-bond between the pyridyl nitrogen atom of the substrates and the formamide group is expected to account for this regioselectivity. Not surprisingly, the C-4 reactivity of 2-chloroquinoline was dramatically diminished and resulted in a low yield (**4h**, 15%

yield). Interestingly, the reaction with 4-chloroquinazoline (contains two electron-withdrawing groups, the pyridine nitrogen atom and the chlorine atom) led to an excellent yield (99% NMR yield after 6 h) even with 0.1 equivalent of benzaldehyde, as shown in Table 3b. In contrast, a yield of 13% was achieved in the absence of the aldehyde. These results suggested that benzaldehyde might function through a catalytic cycle in this reaction. Carboline, a widely used chemical scaffold in drug discovery,¹⁵ was also amendable to this protocol, albeit in moderate yield (**4i** in Table 3a, 50% yield). It is worth noting that this protocol can also be implemented in the synthesis of **4a**, a key intermediate for the preparation of an NMDA antagonist, CGS 19755 (**4a**, 67% yield). The literature yield was 30–40% with the reaction conditions H₂SO₄, 30% H₂O₂, and FeSO₄·7H₂O.²³

Reactions between benzothiazole and *N*-alkylated amides were also investigated (Table 4). *N*-Methylformamide

Table 3. Benzaldehyde-mediated photoredox CDC reactions of six-membered nitrogen-containing heteroaromatics with formamide^a

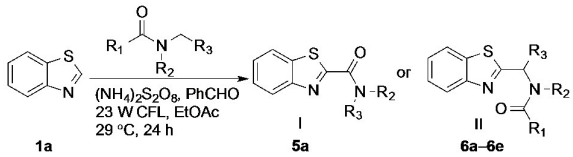


^a See general procedure A for the experimental details unless otherwise noted. Isolated yield was reported for each reaction. ^b Performed without TsOH·H₂O (general procedure B). ^c The reaction to generate **4c**, C-6:C-2:C-4 = 16.5:5.5:1 (r.r.); the reaction to generate **4e**, C-2:C-4 = 3:1 (r.r.); and the reaction to generate **4f**, C-1 product only. The regiometric ratio (r.r.) was determined by the ¹H NMR spectroscopy.

underwent an arylation with approximately 8:1 regioselectivity to activate the carbonyl C_{sp}²-H bond (**5a**, 88% yield). *N,N*-Dimethylformamide (DMF) mainly led to the activation of the C_{sp}³-H bond of the *N*-methyl group, affording an opposite

regioselectivity (**6a**, 80% yield). The different BDEs between carbonyl C_{sp2}-H and *N*-methyl C_{sp3}-H bonds of DMF account for this selectivity.¹⁶ *N*-Methylacetamide and *N,N*-dimethylacetamide (DMA) do not have a C_{sp2}-H bond and only undergo the arylation at the *N*-methyl group by activating the C_{sp3}-H bond (**6b** and **6c**). The reaction with *N*-methyl pyrrolidin-2-one also displayed excellent efficiency (**6d**, 89% yield). The regioselectivity (r.r. = 12:1) favors the methylene carbon over the *N*-methyl carbon. It should be noted that all of these reactions were conducted in the absence of TsOH•H₂O. The reaction with *N,N*-diethylacetamide activated the C_{sp3}-H bond next to the nitrogen atom. However, the yield was low even with the use of one equivalent of TsOH•H₂O (**6e**, 25% yield).

Table 4. Benzaldehyde-mediated photoredox CDC reactions between benzothiazole and different amides^a



Entry	Substrate	Product	Yield
1			88%
2			80%
3			82%
4			87%
5			89%
6			25%

^a See general procedure C for the experimental details. All reactions were reported with the isolated yields unless otherwise noted. ^b Entry 1, 8:1 (r.r) = structure I:structure II; entry 2, 1:9 (r.r) = structure I:structure II; and entry 4, 12:1 (r.r) = 1-substituted methyl pyrrolidin-2-one:1-methyl-5-substituted pyrrolidine-2-one. The regiomer ratio (r.r.) was determined by the ¹H NMR spectroscopy. ^c Performed with TsOH•H₂O (1 equiv).

Mechanism. To further disclose the photochemical nature of the reaction, we performed several control experiments by varying the wavelengths of the light (Fig. 2a). The initial results indicated that the near UV region of the CFL emission spectrum (364 nm) contributed to this conversion (entry 3, no reaction in darkness; entry 4, 60% yield, irradiated with household CFL bulb; and entry 5, 99% yield, irradiated with a black light CFL bulb, UV-A light). A similar yield was achieved in entries 6 and 7 indicated benzaldehyde-enhanced reactivity

was completely suppressed when the reaction was performed under filtered light (>400 nm). The UV-visible absorption spectra of the reagents were collected as shown in Fig. 2b. No obvious absorption was detected for formamide, ethyl acetate, APS and benzothiazole in the near UV region of the CFL emission spectrum (e.g. $\lambda = 364$ nm), while benzaldehyde displayed a strong concentration-dependent absorption (Fig. S3). We then investigated the correlation of the photochemical activity of the reaction with benzaldehyde. No ground state association was detected between the reaction components, as shown in Fig. S4. These experimental results again suggested that benzaldehyde works via an excited state through an $n-\pi^*$ transition.^{17a} However, it is not clear if the triplet state, which plays a dominant role in the photochemistry of benzaldehyde along with a relatively long lifetime, is involved in this process, because the triplet state of benzaldehyde tends to be quenched by acid.^{17a,18}

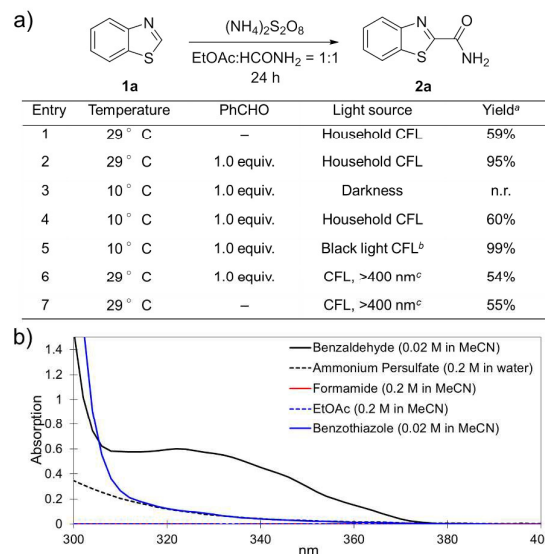


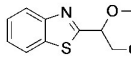



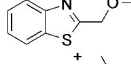
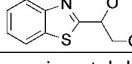


Fig. 2 (a) Investigation of the role of light in the reaction. ^a Yields were determined by ¹H NMR spectroscopy using CH₂Br₂ as an internal standard; ^b UV-A light; ^c water bath with 1 M NaNO₂ aqueous solution to suppress light with $\lambda \leq 400$ nm; and n.r., no reaction. (b) UV-Vis absorption spectra of the reagents.

Coupling reactions of dioxane with benzothiazole were then performed to further investigate the mechanism for the generation of the sulphate radical. Further, we envisioned that this CDC reaction might be used to install an ether functionality to heteroaromatics due to the stability of the C_α radical to an oxygen atom. As shown in Table 5, in the absence of acetamide (please note that acetamide does not have a C–H bond α to the nitrogen atom) and benzaldehyde, no product was observed in the reaction with 1,4-dioxane. A yield of 32% was achieved with the use of 20 equivalents of acetamide. This reaction again suggested that the amide-enhanced decomposition of APS at low temperature accounted for the

reactivity due to hydrogen bonding and donor-acceptor interactions between acetamide and APS.¹² The addition of one equivalent of benzaldehyde to this acetamide-containing reaction mixture improved the yield to 68%. The similar yield was achieved in the absence of acetamide (entry 4 in Table 5, 70% yield). The isolated yield of this reaction was 62%. These results indicated that the photoexcited benzaldehyde worked via direct promotion of APS decomposition and contributed to the initiation of the radical chain reaction. The same result was also obtained for glycol dimethyl ether with the isolated yield of 64% (entry 5 in Table 5).

Table 5. Investigation of the CDC reactions of ethers with benzothiazole^a

Entry	ether	CH ₃ CONH ₂	PhCHO	Product	Yield ^b
1		—	—	—	n.r.
2		20 equiv.	—		32%
3		20 equiv.	1 equiv.	7a	68%
4		—	1 equiv.	7a	70% (62% ^c)
5		—	1 equiv.	 	(64% ^c) r.r = 1:1

^a See general procedure D for the experimental details. ^b Yields were determined by the ¹H NMR spectroscopy using CH₂Br₂ as an internal standard; n.r., no reaction; r.r. = regioselectivity ratio between **7b** and **7c**. ^c Isolated yield.

Substituted benzaldehydes were then used to further investigate the mechanism. 2-Methylbutanal without an aromatic system cannot improve the yield of the reaction (entry 1 in Fig. 3a, 57% yield), which again suggests that benzaldehyde functions through absorbing near UV light from the CFL bulb. No hydrogen/deuterium exchange was observed for the reaction with benzaldehyde- α -d₁ (entry 3, 96% yield), suggesting that benzoyl radical was not involved in this reaction. No yield improvement was observed for the benzaldehyde with strong electron-donating groups, such as 4-methoxybenzaldehyde (entry 4, 63% yield), which was previously reported to promote the atom-transfer radical addition of alkenes as an excellent energy transfer catalyst.^{17a} This result is indicative that benzaldehyde might not work via an energy-transfer pathway in the studied reactions. It should be noted that 4-methoxybenzaldehyde was fully recovered in this reaction. On the other hand, the benzaldehydes with electron-withdrawing groups and weak electron-donating groups promoted the reactions accompanied with partial

decomposition (entries 2 and 5–11, 83–99% yields with 7–71% recovered yields for benzaldehydes). The results suggest that a study of the decomposition of benzaldehyde may help disclose the reaction mechanism. The UV-Vis absorption spectra of different benzaldehydes were collected. As shown in Figure 3b, the substituted benzaldehydes, which displayed stronger absorption in the near UV region of the CFL emission spectrum (e.g. $\lambda = 364$ nm), suffered more decomposition.

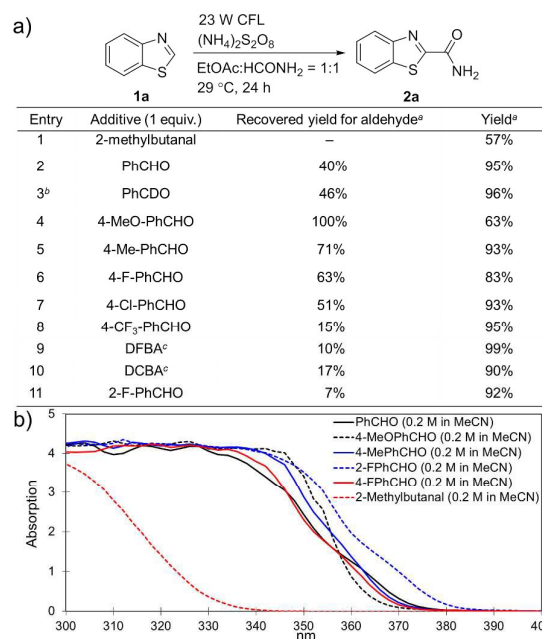
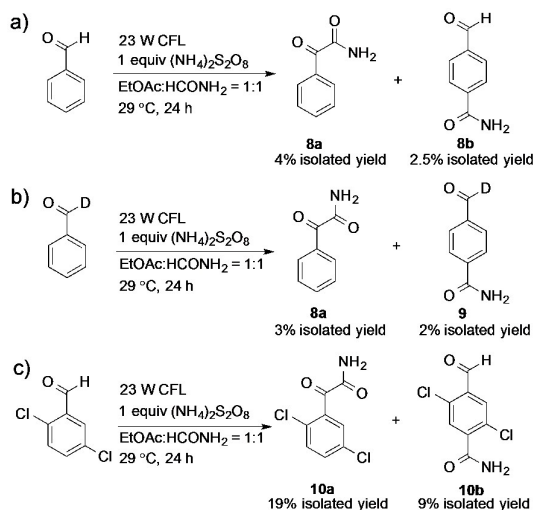


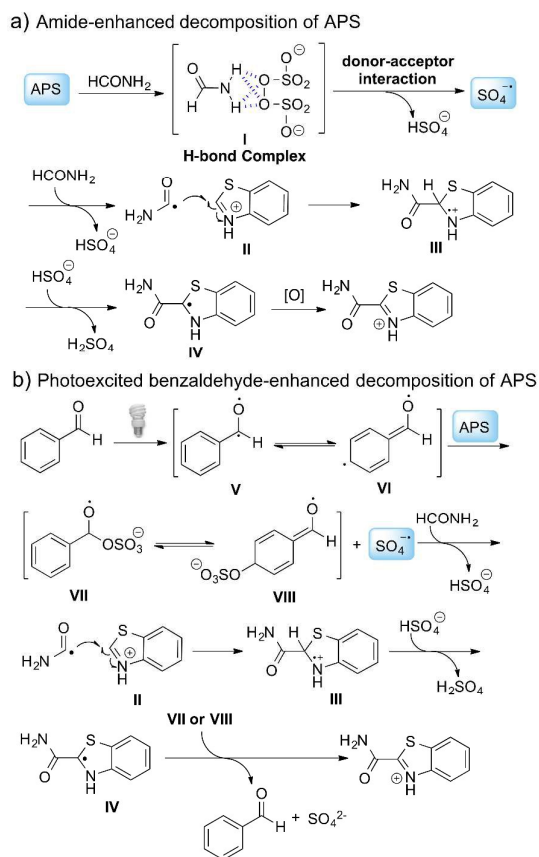
Fig. 3 Investigation of the role of benzaldehyde in the reaction. (a) Evaluation of different aldehydes. ^a Yields were determined by ¹H NMR spectroscopy using CH₂Br₂ as an internal standard; ^b PhCDO was recovered without deuterium exchange; ^c DFBA stands for 2,5-difluorobenzaldehyde, and DCBA stands for 2,5-dichlorobenzaldehyde. (b) UV-Vis absorption spectra for different benzaldehydes.

A control experiment without the addition of the heteroaromatics, such as benzothiazole, was conducted to study the decomposition of benzaldehyde. Only 2-oxo-2-phenylacetamide (**8a**, 4% isolated yield) and 4-formylbenzamide (**8b**, 2.5% isolated yield) were detected, as shown in Scheme 1a. A possible mechanism for the decomposition of benzaldehyde in this reaction is shown in Scheme S1. Photoexcited benzaldehydes **V** and **VI** that are generated upon illumination might be the key intermediates. The coupling reaction between the photoexcited benzaldehydes and the carbamoyl radical, followed by the extraction of hydrogen, could offer **8a** and **8b**. Two other control reactions with benzaldehyde- α -d₁ and 2,5-dichlorobenzaldehyde were also conducted to study the decomposition of the benzaldehydes. Similar results were obtained, as shown in Scheme 1b and 1c.



Scheme 1. Reactions to study the products of benzaldehyde-mediated photoredox CDC reactions.

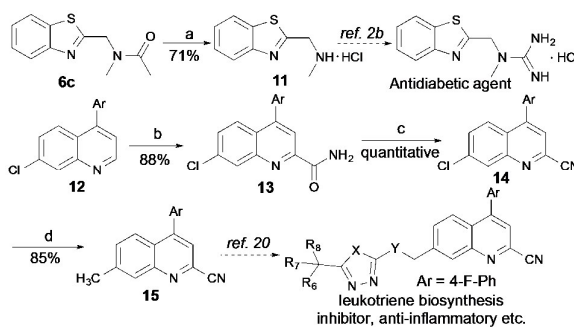
The proposed mechanism is shown in Scheme 2. In the absence of benzaldehyde, hydrogen bonding between APS and



Scheme 2. Proposed reaction mechanism.

formamide leads to the development of complex **I**,¹² followed by the generation of the sulfate radical through the donor-acceptor interaction and the decomposition of this complex. When benzaldehyde is added to the reaction mixture, photoexcited benzaldehydes **V** and **VI** are generated upon illumination¹³ and promote the decomposition of APS to generate the sulfate radical. The carbamoyl radical from the reaction between the sulfate radical and formamide undergoes a nucleophilic addition with the protonated benzothiazole at the C-2 position, followed by deprotonation and oxidation to provide the desired product. It should be noted that intermediates **VII** or **VIII** might be involved in the oxidation step, and thereby regenerated benzaldehyde to finish the catalytic cycle based on the observation that a catalytic amount of benzaldehyde can also promote the reaction (Table 3B). Further validation of the mechanism is still ongoing in our lab.

Application. Many therapeutic drugs feature an amine or ether C_{α} to a heteroarene ring (Fig. S1). The new CDC protocol described in this study is readily applicable to the efficient assembly of compounds in this class (Schemes 3, S2, and S3). For example, starting with **6c**, the hydrolysis reaction led to key intermediate **11** for the synthesis of an antidiabetic agent.^{2b} The new reaction was also successfully applied to the installation of a cyano group to the C-2 position of **12** with excellent yield (88% yield over two steps) through a coupling reaction with formamide followed by a dehydration reaction under moderate reaction conditions, thereby facilitating the synthesis of a leukotriene biosynthesis inhibitor.¹⁹ Given the prevalence of nitrile-containing pharmaceuticals²⁰ and the difficulty to perform the late-stage functionalization for the nitrile group (the previous reactions typically need metal catalysts, high temperature, toxic reagents, etc.),²¹ this new strategy will find broad application in drug discovery programs. In combination with the Suzuki-Miyaura coupling reaction, a new synthetic route was successfully implemented to generate key intermediate **15** with 75% overall yield.



Scheme 3. New CDC strategy to synthesize two key intermediates for biologically active compounds. Reagents and conditions: a) HCl (6 M), EtOH, reflux, overnight; b) 23 W CFL, $(\text{NH}_4)_2\text{S}_2\text{O}_8$, PhCHO, HCONH₂, EtOAc; 29 °C, 24 h; c) TFAA, TEA, NH₂, THF, 0 °C, 2 h; d) Pd(OAc)₂, MeB(OH)₂, SPhos, K₃PO₄, toluene, reflux, 36 h.

Conclusions

In this study, we developed an effective photochemically induced CDC strategy for the α -heteroarylation of amides (α to nitrogen) and ethers. A new decomposition mechanism of APS enhanced by amides and photoexcited benzaldehydes was disclosed. Remarkably, both the C_{sp^3} -H activation (the N -alkyl C-H bonds of the amide or the C_{α} -H bonds of the ether) and the C_{sp^2} -H activation (the carbonyl C-H bond of formamide) were amendable to this protocol. The reaction is environmentally benign and metal-free, is conducted at room temperature with readily available reagents, and uses household CFL bulbs as the light source. These features will likely render this reaction a useful tool for the synthesis of biologically active compounds and the design of new radical-based reactions. However, a large excess of amides and ethers are required for the effective conversion of this reaction. Future research aims to optimize the reaction to decrease the amount of the amide or ether coupling partners and the application of this new strategy to new biologically active compounds.

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Notes and references

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