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Specifically, the decay constant of the cluster is 0.04 Å⁻¹ in 1,2,4-trichlorobenzene (TCB). We consider two possible effects of ligand length on conductance decay with apparent molecular length, we connect these same clusters to incorporating metal chalcogenide molecular clusters in single-molecule electrical circuits. Moreover, we have found that the decay constant is impacted greatly by changing the solvent from molecular size. We recently reported that we could incorporate metal chalcogenide molecular clusters in single-molecule electrical circuits. In this study, in order to determine how transport through such systems depends on molecular length, we connect these same clusters to conducting ligands of varying lengths. We have found that the inclusion of the cluster in the molecular circuit reduces the effect of ligand length on conductance decay with apparent molecular size. Moreover, we have found that the decay constant is impacted greatly by changing the solvent from 1,2,4-trichlorobenzene (TCB) to 1-bromonaphthalene (BrN). Specifically, the decay constant of the cluster is 0.04 Å⁻¹ in BrN, while it is 0.12 Å⁻¹ in TCB. We consider two possible mechanisms to explain these remarkable observations. This work demonstrates, for the first time, a molecular system where the tunneling decay constant can be modified by altering the environment around the molecule.

Introduction

Controlling charge transport through molecular electrodes is critical to the realization of nanoscale electronic devices. While numerous organic molecules have been studied as connecting wires for single-molecule junction studies, very little is known about the effect of metal complexes in these types of junctions. We recently reported that we could incorporate metal chalcogenide molecular clusters in single-molecule electrical circuits. In this study, in order to determine how transport through such systems depends on molecular length, we connect these same clusters to conducting ligands of varying lengths. We have found that the inclusion of the cluster in the molecular circuit reduces the effect of ligand length on conductance decay with apparent molecular size. Moreover, we have found that the decay constant is impacted greatly by changing the solvent from 1,2,4-trichlorobenzene (TCB) to 1-bromonaphthalene (BrN).

Results and Discussion

The single cluster circuits that we have designed, assembled, and studied consist of an atomically defined Co₂Se₇ molecular cluster⁶, ¹⁷ (Fig. 1a) wired between nanoscale electrodes. The wiring is formed from bifunctional, conjugated ligands (Fig. 1b) that bind specifically and directionally to the electrode and to the cluster. We employ an atomically defined segment of polyacetylene that has an arylphosphine group on one terminus that coordinates to a cobalt atom on the clusters and an arylthiomethyl group on the other terminus that attaches to the Au electrode. The mono-, di-, and triene ligands are L₁, L₂, and L₃ and the corresponding clusters are 1, 2, and 3, respectively. Fig. 1c shows the molecular structure of 1 as determined by single crystal X-ray diffraction (SCXRD).ⁱ⁵

We measured the conductance of both the individual molecular clusters (1–3) and the free conducting ligands (L₁–L₃) using a scanning tunneling microscope based break-junction (STM-BJ) technique.²¹ In this technique, an Au STM tip and substrate are repeatedly brought into and out of contact to form and break Au-Au point contacts in solutions of the target compounds. During this process, a bias voltage is applied across the junction while current is measured in order to determine conductance (G = I/V) of the junction. The measurements are repeated thousands of times, and the data is analysed to reveal statistically significant results. The data is processed by compiling thousands of individual conductance traces into one-dimensional, logarithmically-binned conductance histograms. We further generate two-dimensional (2D) histograms of the conductance versus displacement by aligning each conductance trace after the
In order to characterize transport through the molecular clusters and the free ligands, we measured the conductance of 1-3 and L1-L3 in two different solvents, BrN and TCB. These solvents were chosen taking into consideration the solubility of both the ligand and the cluster systems as well as for their varied affinity to gold electrodes.\(^{23}\) Fig. 2a and 2b contain the one-dimensional histograms for the measurements in BrN, and Fig. 2d and 2e show the same for the measurements in TCB. The 2D histograms for 1 and L1 in each solvent are an inset in the respective figures. The 2D histograms show a significant difference in length of the molecular feature for 1 and for L1. Moreover, the cluster junction lengths measured from the 2D histograms correlate with the molecular lengths of the cluster with the ligands fully extended (measured in BrN: 9 Å and 21 Å, and expected from SCXRD: 13 Å and 32 Å, for L1 and 1 respectively). Despite the additional complexity of the cluster system, we conclude that we are indeed probing transport through Au-ligand-cluster-ligand-Au junctions based on this large difference in the observed lengths. Furthermore, the histograms in Fig. 2a and 2d show shoulders, with an increasing prominence for the longer-ligand systems. Comparing the conductance of these shoulders with the ligand conductance in Fig. 2b and 2e, we attribute these shoulders to free ligands, that is, ligands that have detached from the clusters.

We fit the peaks of all conductance histograms for both solvents with a Gaussian function and plot the peak conductance values versus the number of C=C units or “enes” in each molecule (Fig. 2c and 2f). In both solvents, and for both free ligand and cluster, we observe that the conductance decreases exponentially with increasing molecular length following the relationship \(G \sim e^{-\beta n}\), where \(n\) is the number of “ene” units in the backbone and \(\beta\) is the decay constant. We report the decay constant per Angstrom using a length of 2.48 Å per “ene” unit. The decay constant for the free ligand series is essentially independent of the solvent (\(\beta = 0.15 \text{ Å}^{-1}\) in TCB and \(0.17 \text{ Å}^{-1}\) BrN).

The unexpected result is the factor of 3 difference in the decay constants of the cluster series in different solvents as can be seen comparing Figures 2c and 2f. In TCB, the \(\beta\) of the cluster system is 0.12 Å\(^{-1}\), and in BrN it is 0.04 Å\(^{-1}\). We note that the difference between the decay constant of the ligand and that of the cluster is greater in BrN than in TCB. Furthermore, the absolute values of the conductance of the cluster series are significantly higher when measured in BrN than in TCB, with the conductance of 3 being almost an order of magnitude higher in BrN compared to TCB. Such a solvent-induced effect on the conductance has been observed in other systems, and this has been attributed to the solvent’s ability to modulate the electrode work function.\(^{23, 24}\)
charge transport through these junctions. Charge transport can occur via a coherent off-resonance process through an orbital on the ligand-cluster-ligand assembly that is coupled to both electrodes. In that situation, the conductance depends on at least two related factors: (1) the energy of this conducting orbital relative to the metal $E_F$, and (2) the coupling between this orbital and both electrodes. As the length of the molecule increases, the HOMO-LUMO gap narrows, and if conductance were just related to energy level alignment, one would naively expect conductance to actually increase. However, transport through the junction is also related to how well the conducting orbital overlaps with the leads, and since the orbital is more delocalized over a longer conjugated molecule, this overlap decreases with increasing length. The conductance thus typically decays exponentially with increasing molecular length. Specifically, as the conjugated backbone gets longer, the molecular orbital is delocalized over a larger molecule, and since the orbital is normalized, a smaller fraction of its amplitude resides on the sulfur atoms; therefore the coupling between the molecule and the electrodes decreases.

If we assume that the conducting orbitals of the cluster and of the ligand for a given length are similar in both character and energy, we can develop a simple tight-binding model of the molecular junctions. The objective of this model is not to reproduce the experimental data but to examine and illustrate how the additional electronic structure of the cluster impacts transport through the system. Our tight-binding model is schematically presented in the insets of Fig. 3a and 3b for the ligand and the cluster respectively. For the conducting ligands, we assign a single energy level, $\epsilon$, for each unit, and allow nearest neighbours to be coupled by $\delta$. The terminal units are coupled to the Au electrodes using an imaginary self-energy, $\iota\gamma/2$. We apply a similar model for the cluster, adding an additional energy level, $E_0$, between two ligands and coupling this site to its nearest neighbour ligand states with $\tau$. We computed the transmission functions for these model systems using a Green’s function approach (see the Supporting Information for a detailed description).

Sample computed transmission functions are shown in Fig. 3a and 3b using the same values for $\epsilon$, $\delta$ and $\tau$ for the ligand and the cluster series. The transmission functions display resonances at energy values corresponding to the molecular orbitals of the system where the probability of an electron being transmitted through the system is unity. The transmission function for L1 contains one resonance at energy $\epsilon$, while longer molecules have resonances equal to the number of sites in the corresponding model. As the length of the molecule increases, the frontier resonance moves closer to $E_F$ but also narrows, which is a consequence of the frontier orbital being delocalized over a longer molecular backbone. Upon comparing the transmission functions for the ligands with those of the clusters, we see that the clusters contain resonances that are closer to $E_F$ than their ligand counterparts, but with narrower resonances. This observation leads to a lower transmission at $E_F$ to which the experimentally measured values should correspond; more importantly, it also leads to a conductance that is more sensitive to the exact location of the $E_F$.

In Fig. 3c, we show the conductances that are determined from the tight-binding model for each molecule versus the number of “ene” units in the molecule (using the same model parameter values for both series). From the fit to these values, it is clear that the predicted decay constants are essentially the same for the ligand and cluster series. We use one set of $E_0$ and $\tau$ values to calculate the representative transmission/conductance functions shown in Fig. 3a and 3b. Regardless of what value is assigned to $E_0$ and $\tau$, we find that this model predicts very similar decay constants for the two systems (Fig. 3d).

Our tight-binding calculations suggest that the addition of a cluster level $E_0$ between two ligands cannot explain the observed change in $\beta$. In other words, the ligand and cluster series should have the same $\beta$ values, unless the energy alignment of the cluster resonance is altered relative to the electrode Fermi level in this model. We have three sets of observations that are consistent with a change in $E_F$: 1) $\beta$ of the cluster in BrN is significantly lower than in TCB, 2) $\beta$ values measured in both solvents are almost the same for the ligand series, and 3) $\beta$ of the cluster is lower than that of the ligand in both solvents. The steeper transmission curves of the cluster series in Fig. 3 indicate that the resonance energies are closer to $E_F$. Within this coherent transport model, we can see that a small change in $E_F$ will result in a large shift in $\beta$ for the cluster relative to the ligand. For instance, changing $E_F$ by -0.5 eV...
shifts the $\theta$ value to 0.1 $\text{Å}^{-1}$ for the clusters while a similar change in $E_F$ for the ligand changes $\theta$ to 0.3 $\text{Å}^{-1}$. These results, when viewed in light of the known ability of solvent-binding to produce changes in $E_F$, point to BrN shifting $E_F$ closer to resonance relative to TCB. This effect is compounded by the sensitivity of the metal. The free ligand and the cluster have very different characteristics (e.g., size, steric hindrance, redox behaviour, dielectric constant polarizability and binding ability) that will result in different shifts in $E_F$.

We also consider a hopping mechanism for charge transfer, a process generally mediated by an activation-controlled reaction (e.g., a thermally induced conformational change or an electron transfer reaction).\cite{1,29} We first rule out the possibility that such a conformational change can occur within the ligand.\cite{8} We also refute the process involving a direct through-space charge transfer from the electrode to an unoccupied molecular level on the cluster through a resonant transfer process.\cite{9} In this picture, the cluster does not have to be chemically attached to the electrodes to form a conducting junction and the charge transfer efficiency depends on the core-electrode spacing. We discount this mechanism based on a previously published study in which we demonstrated that our clusters form molecular junctions by bonding their terminal thiomethyl groups to the Au electrodes.\cite{15} By varying the substitution pattern or removing the aurophilic functionality, we can modulate or completely shut down the conductivity of these molecular junctions, suggesting that there is an orbital pathway for the transport of charge in these cluster systems. These findings refute the idea of direct through-space charge transfer mediated by an orbital localized on the core.

We are left with a hopping mechanism in which the charge tunnels from the source electrode across the ligand to the cluster core and then transfers to the drain electrode through a second coherent tunneling process. Such a transport process requires that the cluster can reversibly change its oxidation state with each charge transfer. Since the applied bias in these measurements is not small (~ 0.5 V) and the cluster core state with each charge transfer. Since the applied bias in these Co clusters can reversibly change its oxidation state, we can modulate or completely shut down the conductivity of these molecular junctions, suggesting that through-space charge transfer mediated by an orbital localized on the cluster core and then transfers to the drain electrode through a second coherent tunneling process. Such a transport process requires that the cluster can reversibly change its oxidation state with each charge transfer. Since the applied bias in these measurements is not small (~ 0.5 V) and the cluster core state with each charge transfer. Since the applied bias in these Co clusters can reversibly change its oxidation state, we can modulate or completely shut down the conductivity of these molecular junctions, suggesting that through-space charge transfer mediated by an orbital localized on the cluster core.

Conclusions

In summary, we measured charge transport through molecular clusters with ligands of different lengths and showed that the conductance decay depends on the solvent used for these measurements. Our results illustrate a novel effect that allows the environment to alter the conductance decay constants. This study opens up the possibility to carry out conductance measurements in which clusters can be controllably gated by changing the environment.\cite{16} While the conducting ligands alone are limited to a one-dimensional system, the three-dimensional architecture of the metal chalcogenide cluster allows us to envision novel electronic devices where a molecular cluster is contacted by electrodes at multiple locations.

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References

29. We note that temperature dependent measurements using an STM-based break-junction experimental setup cannot distinguish a hopping mechanism from an off-resonant transport mechanism as simple molecules such as 1,4-diaminobenzene show a strong temperature dependent conductance (M. Kamenetska, Ph.D. Thesis, Columbia University, 2012.)