

# Photochemical & Photobiological Sciences

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## Photochemical processes induced by the irradiation of 4-hydroxybenzophenone in different solvents

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The singlet and triplet excited states of 4-hydroxybenzophenone (4BPOH) undergo deprotonation in the presence of water to produce the anionic ground-state, causing fluorescence quenching and photoactivity inhibition. The same process does not take place in an aprotic solvent such as acetonitrile. In acetonitrile, 4BPOH is fluorescent (interestingly, one of its fluorescence peaks overlaps with peak C of humic substances), it yields singlet oxygen upon irradiation and induces the triplet-sensitised transformation of phenol (with a rate constant of  $(6.6 \pm 0.3) \cdot 10^7 \text{ M}^{-1} \text{ s}^{-1}$  ( $\mu \pm \sigma$ ) between phenol itself and triplet 4BPOH). The 4BPOH shows an intermediate behaviour in a partially protic solvent such as 2-propanol, where some deprotonation of the excited states is observed. In acetonitrile/2-propanol mixtures (at least up to 50% of 2-propanol) there is also some evidence of alcohol oxidation by the 4BPOH triplet state, while the experimental data are silent concerning such a possibility in pure 2-propanol. Considering that benzophenones are important components of chromophoric dissolved organic matter (CDOM) in surface waters, the present findings could have significance for the photoactivity of the hydrophilic surface layers vs. the hydrophobic cores of CDOM.

## Introduction

Benzophenones are important chromophores and photosensitisers contained in chromophoric dissolved organic matter (CDOM) in natural waters, and particularly in humic substances.<sup>1,2</sup> Recent evidence has provided additional arguments in favour of an important role played by benzophenones in the optical and photochemical properties of CDOM and of humic material.<sup>3,4</sup>

A major feature of benzophenones is the fact that the energy levels of their first excited singlet ( $S_1$ ) and triplet ( $T_1$ ) states are very near. This issue, combined with the fact that their  $S_1$  and  $T_1$  are strongly coupled by spin-orbit interactions, favours the inter-system crossing (ISC) from  $S_1$  to  $T_1$  that has elevated (often near-unity) quantum yields. Therefore, light absorption exciting electrons from  $S_0$  to  $S_1$  is followed to a large degree by relaxation of the excited

benzophenones into  $T_1$ , which triggers important triplet-sensitised processes.<sup>5</sup> In the case of unsubstituted benzophenone (BP), both  $S_1$  and  $T_1$  have  $n-\pi^*$  configuration, independently of the polarity of the solvent. The  $n-\pi^*$  configuration favours both the ISC and the reactivity of  $T_1$ , thereby enhancing the triplet-sensitised processes. In the case of substituted benzophenones, the  $n-\pi^*$  and  $\pi-\pi^*$  configurations in both the singlet and triplet manifolds are relatively near in energy to each other. Moreover their relative positions depend on the polarity of the solvent, and the prevailing configuration may be different in solvents of differing polarity. Interestingly, the  $\pi-\pi^*$  configurations of both  $S_1$  and  $T_1$  are associated to longer excited-state lifetimes, lower ISC quantum yields and reduced photochemical reactivity.<sup>5-9</sup>

In the case of 4-hydroxybenzophenone (4BPOH) the  $S_1$  state has  $n-\pi^*$  configuration, independently of the polarity of the solvent. In contrast, the  $n-\pi^*$  and  $\pi-\pi^*$  configurations of  $T_1$  are relatively near in energy. Apolar solvents favour  $n-\pi^*$ , while polar ones give higher stability to  $\pi-\pi^*$ . In some cases one can even observe that  $T_1$  has a mixed  $n-\pi^*$  /  $\pi-\pi^*$  character, because of very similar energy levels of the two configurations, which may produce a double-exponential decay of the phosphorescence signal.<sup>5</sup> In acetonitrile, an equilibrium between the  $n-\pi^*$  and  $\pi-\pi^*$  triplet states of 4BPOH has been reported.<sup>6</sup>

The described effect of polarity on the excited states of 4BPOH is important in aprotic solvents, while the picture may change dramatically in protic ones such as water. Indeed, 4BPOH- $S_0$  is a weak acid ( $pK_a \sim 8.5$ ), while 4BPOH- $S_1$  and 4BPOH- $T_1$  are strong acids that undergo fast deprotonation/deactivation in an aqueous

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environment.<sup>10</sup> This issue strongly limits the photoinduced reactivity of 4BPOH in protic solvents compared to aprotic (polar or apolar) ones.<sup>11</sup>

The fluorescence spectrum of 4BPOH is also interesting, because it shows a band in a region that overlaps with peak C of humic substances (excitation at 300–320 nm, emission at 450–550 nm).<sup>12–14</sup> This fact makes 4BPOH a potentially interesting model molecule for humic substances, and it accounts for the importance of studying and correlating its photophysics and photochemistry in different solvents.

Humic substances are a complex mixture of compounds that originate from a wide variety of sources.<sup>15,16</sup> Their components are assembled in a micelle-like fashion where hydrophilic moieties are in contact with water, while the hydrophobic ones form waterless inner cores.<sup>17,18</sup> Humic compounds are also important sunlight absorbers,<sup>19</sup> with key consequences for the penetration of biologically harmful UV radiation in the water column.<sup>20,21</sup> Sunlight absorption produces a variety of reactive transient species (triplet states, singlet oxygen <sup>1</sup>O<sub>2</sub> and hydroxyl radical •OH)<sup>22–25</sup> that are involved in the transformation of dissolved pollutants.<sup>26–28</sup> The photoactivity of humic substances is linked to their molecular weight,<sup>29–31</sup> because the larger fractions undergo important intermolecular interactions that favour internal conversion and quench both photoactivity and fluorescence.<sup>4,32–35</sup> However, the hydrophobic cores of the larger moieties show elevated concentration values of singlet oxygen.<sup>36–38</sup>

Because 4BPOH shares some features in common with humic compounds, which make it a potentially interesting model molecule, the present work investigates the photochemical processes induced by 4BPOH in different solvents, namely an aprotic one (acetonitrile), a partially protic one (2-propanol), as well as water. The results may apply to photoinduced reactions taking place in aqueous solution or in hydrophobic environments, such as those occurring in different CDOM compartments.<sup>17,18</sup>

## Experimental

**Reagents and materials.** 4-Hydroxybenzophenone (4BPOH, 98%), benzophenone (BP, 99%), furfuryl alcohol (FFA, 98%), 2-propanol (99.8%), 2-nitrobenzaldehyde (98%), acetonitrile (gradient grade), HClO<sub>4</sub> (70% in water), NaOH (98%) and methanol (gradient grade) were purchased from Sigma-Aldrich, nitrogen (5.5, 99.9995%), argon (5.5, 99.9995%) and oxygen (6.0, 99.9999%) from SAPIO. Water used was of Milli-Q quality.

**Fluorescence measurements.** The excitation-emission matrix (EEM) fluorescence spectra were taken with a VARIAN Cary Eclipse Fluorescence Spectrophotometer, with an excitation range from 200 to 400 nm at 10 nm steps, and an emission range from 200 to 600 nm with a scan rate of 1200 nm min<sup>-1</sup>. Excitation and emission slits were both set at 10 nm. Spectra were taken in a fluorescence quartz cuvette (Hellma) with 1 cm optical path length. The Raman signal of water was taken as a reference for lamp intensity and signal stability within different measurements.

**Laser flash photolysis experiments.** The laser apparatus has been previously described.<sup>39</sup> The laser excitation wavelength was fixed at

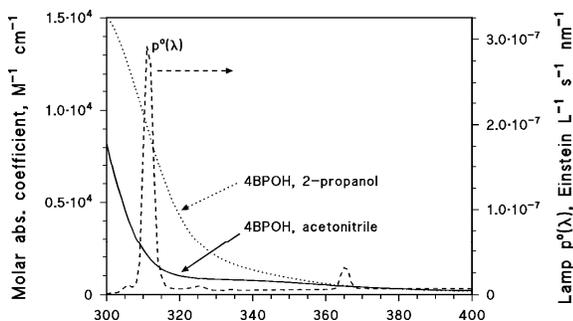
355 nm with an energy of ~94 mJ/pulse. An appropriate volume of stock solutions was mixed before each experiment to obtain the desired concentrations. The pH was acidified using HClO<sub>4</sub>, and deoxygenated and oxygen-saturated solutions were used after 20 min of bubbling with argon or pure oxygen, respectively.

The second-order rate constant between 4BPOH-T<sub>1</sub> (hereafter, <sup>3</sup>4BPOH\*) and quenchers was calculated from the regression line of the absorbance logarithm decay against quencher concentration. The error bars were derived at the 3σ level from the scattering of the experimental data.

**Steady irradiation experiments.** Solutions (5 mL volume) containing 0.1 mM 4BPOH and other reagents (when necessary, such as 0.1 mM phenol or 0.1 mM FFA) were placed in cylindrical Pyrex glass cells (4.0 cm diameter, 2.3 cm height, with a lateral neck tightly closed with a screw cap). The solvent was either water, acetonitrile or 2-propanol. Irradiation of the cells took place under a Philips TL 01 (20W) lamp with emission maximum at 313 nm. The solutions in the cells were irradiated from the top and the optical path length was 0.4 cm. The lamp irradiance over the solutions was 2.6±0.1 W m<sup>-2</sup>, measured with a CO.FO.ME.GRA. (Milan, Italy) power meter equipped with a UV-sensitive probe (290–400 nm). The lamp emission spectra (spectral photon flux density  $p^\circ(\lambda)$ , units of Einstein L<sup>-1</sup> s<sup>-1</sup> nm<sup>-1</sup>) were obtained by combining spectrophotometric measures (CCD spectrophotometer Ocean Optics USB 2000, calibrated with a DH-2000-CAL radiation source) with chemical actinometry using 2-nitrobenzaldehyde. The detailed procedure to determine  $p^\circ(\lambda)$  is described in detail elsewhere.<sup>25</sup> Figure 1 reports the spectral photon flux density of the lamp, together with the absorption spectra of 4BPOH (molar absorption coefficients) in both acetonitrile and 2-propanol. The latter were taken with a Varian Cary 100 Scan double-beam UV-vis spectrophotometer, using quartz cuvettes (Hellma, optical path length 1 cm). The main reason for the higher absorption coefficient of 4BPOH in 2-propanol above 300 nm is that the absorption maximum in this solvent, compared to acetonitrile, is shifted from 285 to 295 nm (bathochromic shift). In contrast, the absorbance at the maximum is similar in both cases.

After the scheduled irradiation times, the cells were withdrawn from the lamp. Measured aliquots of the irradiated solutions (2 mL) were diluted with an equal volume of water and analysed by high-performance liquid chromatography with diode-array detection (HPLC-DAD). It was used a VWR-Hitachi Elite instrument, equipped with L-2300 autosampler (25 μL injection volume), L-2130 quaternary pump for low-pressure gradients, Duratec DDG-75 online degasser, L-2300 column oven (operated at 40 °C), L-2445 DAD detector and column RP-C18 LiChroCART packed with LiChrospher 100 RP-18 (VWR, 4 mm × 125 mm × 5 μm).

The eluent was a mixture of 3.5 mM H<sub>3</sub>PO<sub>4</sub> in water (A) and of methanol (B). Gradient elution (1 mL min<sup>-1</sup> flow rate) was as follows: from 5% B to 70% B in 15 min, keep at 70% B for 7 min, back to 5% B in 5 min and keep at 5% B for an additional 5 min (post-run equilibration). The detection wavelengths were 290 nm for 4BPOH, 210 nm for phenol, and 230 nm for FFA. The retention times were 15.5 min for 4BPOH, 9.0 min for phenol and 5.3 min for FFA. The column dead time was 1.4 min.



**Figure 1.** Absorption spectra (molar absorption coefficients) of 4BPOH in acetonitrile and 2-propanol. Emission spectrum (spectral photon flux density) of the lamp used for the steady irradiation experiments.

Additional runs were carried out to measure the formation of 4-phenoxyphenol (4PP) from 1 mM phenol + 1 mM 4BPOH in acetonitrile and 2-propanol. Isocratic elution (1 mL min<sup>-1</sup> flow rate) used 50% A and 50% B, with 229 nm detection wavelength and 10.2 min retention time for 4PP.

The time evolution data of 4BPOH, phenol and FFA were fitted with pseudo-first order kinetic functions of the form  $C_t = C_o e^{-k t}$ , where  $C_t$  is the concentration of the substrate at the time  $t$ ,  $C_o$  its initial concentration and  $k$  the pseudo-first order degradation rate constant. The initial transformation rate was calculated as  $R_o = k C_o$ . The uncertainty on the rates is reported as  $\pm \sigma$  and it mainly depends on the uncertainty on  $k$ , which represents the goodness of the fit of the exponential functions to the experimental data. The reproducibility of replicated experiments was around 15%.

The calculation of the quantum yields of 4BPOH and phenol degradation was based on the photon flux absorbed by the photosensitizer, 4BPOH ( $P_a^{4BPOH}$ ). The latter was calculated as follows:<sup>40</sup>

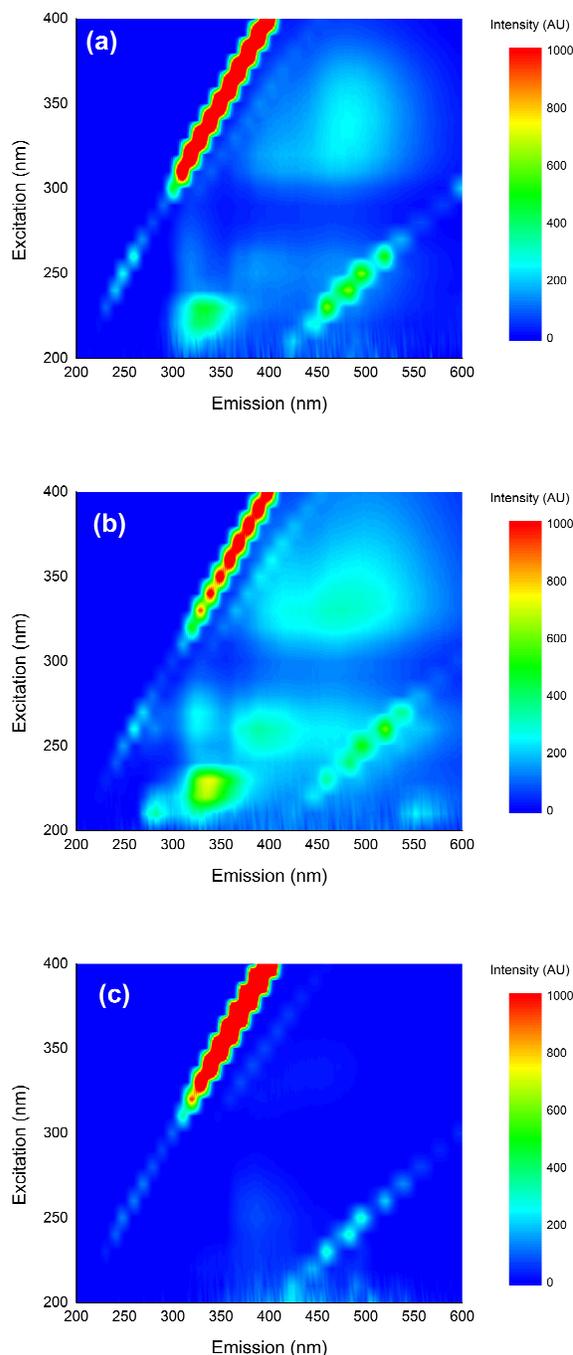
$$P_a^{4BPOH} = \int_{\lambda} p^o(\lambda) [1 - 10^{-\epsilon_{4BPOH}(\lambda) b [4BPOH]}] d\lambda \quad (1)$$

where  $p^o(\lambda)$  is the spectral photon flux density of the lamp,  $\epsilon_{4BPOH}(\lambda)$  the molar absorption coefficient of 4BPOH,  $b = 0.4$  cm the optical path length in solution and  $[4BPOH] = 1 \cdot 10^{-4}$  M the initial concentration of the photosensitizer.

## Results and discussion

### Fluorescence measurements

Figure 2 reports the fluorescence EEM spectra of 0.1 mM 4BPOH in different solvents (acetonitrile, 2-propanol and water, the latter at the natural pH  $\sim 6$ ). Considering first the fluorescence spectra in the organic solvents, one can notice a band at Ex/Em  $\sim 225/325$  nm that is included in the region of phenolic compounds (which is consistent with the OH group on the aromatic ring of 4BPOH).<sup>12</sup>



**Figure 2.** EEM fluorescence spectra of 0.1 mM 4BPOH in acetonitrile (a), 2-propanol (b) and water (c).

More interestingly, the band with Ex/Em  $\sim 325/475$  nm overlaps with peak C of humic substances.<sup>12</sup> Because benzophenones are known chromophores/photosensitisers occurring in CDOM, the fact that a compound of this class shows fluorescence in the humic region deserves a certain attention. In contrast to the results obtained in organic solvents, the fluorescence spectrum of 4BPOH in water shows

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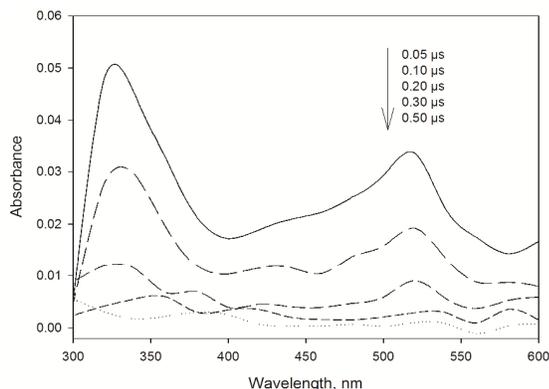
little or no fluorescence emission. This issue is consistent with literature reports that the excited states of 4BPOH (including the excited singlet state) undergo rapid (radiationless) deactivation by deprotonation in aqueous solution.<sup>41,42</sup>

### Laser flash photolysis experiments

The laser excitation of 4BPOH in acetonitrile produces a transient spectrum with absorption maxima at 330 and 520 nm, with the characteristic features of <sup>3</sup>4BPOH\* (see Figure 3).<sup>43</sup> The transient spectra taken at relatively long times (about 0.30 μs) show the occurrence of a small peak between 550 and 600 nm, which might be consistent with the ketyl radical of 4BPOH (4BP(H<sup>•</sup>)OH). The latter could arise upon reaction between <sup>3</sup>4BPOH\* and ground-state 4BPOH (<sup>3</sup>4BPOH\*+4BPOH→4BP(H<sup>•</sup>)OH+4BPO<sup>•</sup>), while the reaction between <sup>3</sup>4BPOH\* and acetonitrile is rather unlikely.<sup>6</sup>

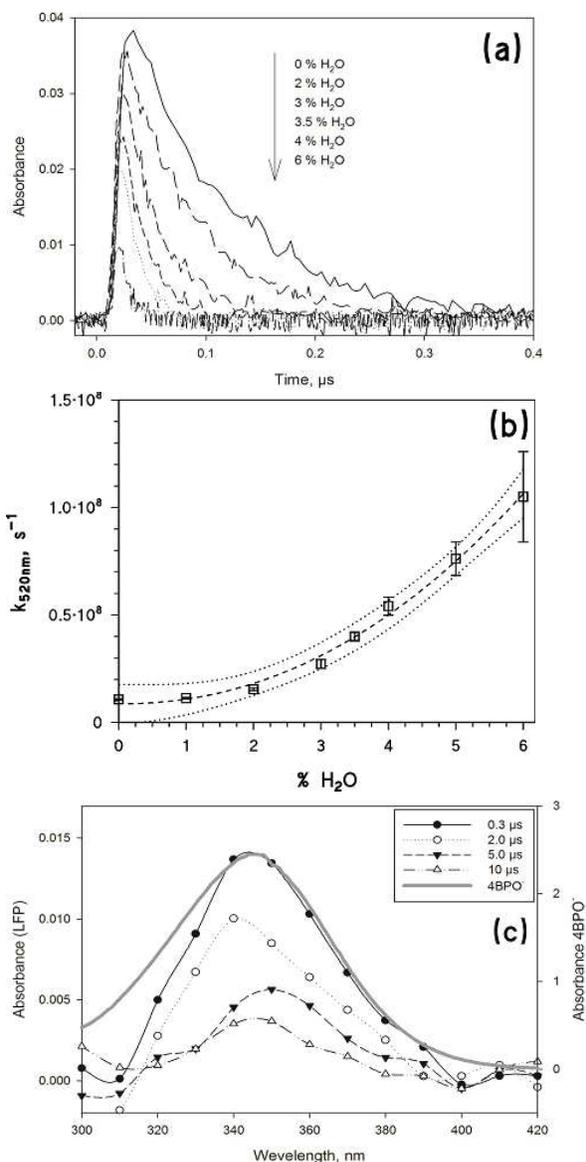
The first-order decay constant of the signal measured at 520 nm ( $k_{520\text{ nm}}$ ) is in the range of  $10^7\text{ s}^{-1}$  in aerated solution. Coherently with the triplet state assignment,  $k_{520\text{ nm}}$  depends on the content of dissolved oxygen. In fact,  $k_{520\text{ nm}}$  increased when passing from an Ar-bubbled system to an aerated and an O<sub>2</sub>-bubbled one. The concentration of dissolved oxygen was assessed based on published data of its solubility in acetonitrile.<sup>44</sup> From the plotted data (see Figure ES1) one gets a second-order reaction rate constant of  $(3.6\pm 0.1)\cdot 10^9\text{ M}^{-1}\text{ s}^{-1}$  between <sup>3</sup>4BPOH\* and O<sub>2</sub>, with the probable formation of <sup>1</sup>O<sub>2</sub> that is typical of triplet-state reactivity.<sup>45</sup>

Figure 4a shows the decay traces of the transient absorbance at 520 nm in water/acetonitrile mixtures, as a function of the volume percentage of water (up to 6%). One can observe that the trace absorbance just after the laser pulse is lower in the presence of higher percentages of water, and that the trace decay becomes faster when increasing the water content. This result can be accounted for by the deprotonation of both the singlet and the triplet states of 4BPOH,<sup>41,42</sup> which would both be favoured in the presence of higher percentages of water. The deprotonation of the singlet state would compete with the formation of the triplet, thereby decreasing the triplet absorbance just after the laser pulse.



**Figure 3.** Time trend of the transient absorption spectrum obtained upon LFP (355 nm, 94 mJ) excitation of 4BPOH (0.1 mM) in aerated CH<sub>3</sub>CN solution at ambient temperature (295 ± 2 K).

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**Figure 4.** (a) Trace decay at 520 nm as a function of the percentage of water in CH<sub>3</sub>CN. The transients were obtained upon LFP irradiation (355 nm, 94 mJ) of 0.1 mM 4BPOH in aerated solutions.

(b) Trend of the first-order decay constant of the absorbance at 520 nm, as a function of the percentage of water in binary mixtures H<sub>2</sub>O/CH<sub>3</sub>CN. 4BPOH 0.1 mM, laser excitation at 355 nm, 94 mJ/pulse. The fit curve (dashed) is a polynomial, the dotted ones are the 95% confidence bands of the fit.

(c) Time trend of the transient absorption spectrum obtained upon LFP (355 nm, 94 mJ) excitation of 4BPOH (0.1 mM) in aerated H<sub>2</sub>O solution at ambient temperature (295 ± 2 K). The grey solid line shows the UV-vis spectrum of anionic 4BPOH (4BPO<sup>-</sup>, 0.1 mM) in water at pH 10.5.

The deprotonation of the triplet state would enhance the relevant trace decay, as reported in Figure 4b which shows that  $k_{520\text{ nm}}$  increases with the volume percentage of water. In the presence of pure water as solvent, the signal corresponding to the triplet state of 4BPOH could no longer be

observed. Most likely,  $^3\text{4BPOH}^*$  is not formed in water because the deprotonation of the excited singlet state is much faster than the ISC. Moreover, even if formed to some extent in water, or in water-acetonitrile mixtures with rather elevated water content,  $^3\text{4BPOH}^*$  is expected (i) to undergo very fast deprotonation, and (ii) to be of  $\pi\text{-}\pi^*$  configuration, which would be strongly stabilised by the polar solvent.<sup>5,6</sup> Both issues would produce negligible photochemical reactivity of  $^3\text{4BPOH}^*$  in the relevant conditions.

The excited-state deprotonation is known to yield an anionic species with absorption maximum around 350 nm.<sup>11,42</sup> In both water and water-acetonitrile mixtures we observed the formation of a transient with maximum absorbance at 350 nm, which was longer-lived compared to the triplet state. The transient absorption spectrum is very similar to that of the anionic form of 4BPOH (hereafter,  $4\text{BPO}^-$ ), as shown in Figure 4c, which justifies the identification of the observed species with ground-state  $4\text{BPO}^-$ .

Interestingly,  $k_{350\text{ nm}}$  (the pseudo-first order decay constant of  $4\text{BPO}^-$ ) increases linearly when increasing the volume percentage of water, which may suggest an acid-base process involving  $4\text{BPO}^-$  (Figure ES12). Considering that the decay of the 350 nm signal brings the absorbance back to its initial value (namely the value observed before the laser pulse, in the presence of 4BPOH alone), the most likely process is the protonation of  $4\text{BPO}^-$  back to 4BPOH. Such a deprotonation-protonation sequence would be accounted for by the fact that the triplet state of 4BPOH ( $\text{pK}_a < -2$ ) is a much stronger acid than the corresponding ground-state ( $\text{pK}_a \sim 8.5$ ).<sup>42</sup> The above hypothesis is further confirmed by the fact that  $k_{350\text{ nm}}$  depends on pH (see Figure ES13). The pH trend would be consistent with a reaction between  $4\text{BPO}^-$  and  $\text{H}_3\text{O}^+$  at  $\text{pH} < 5$ , while at  $\text{pH} > 5$  the prevailing reaction would take place with water.

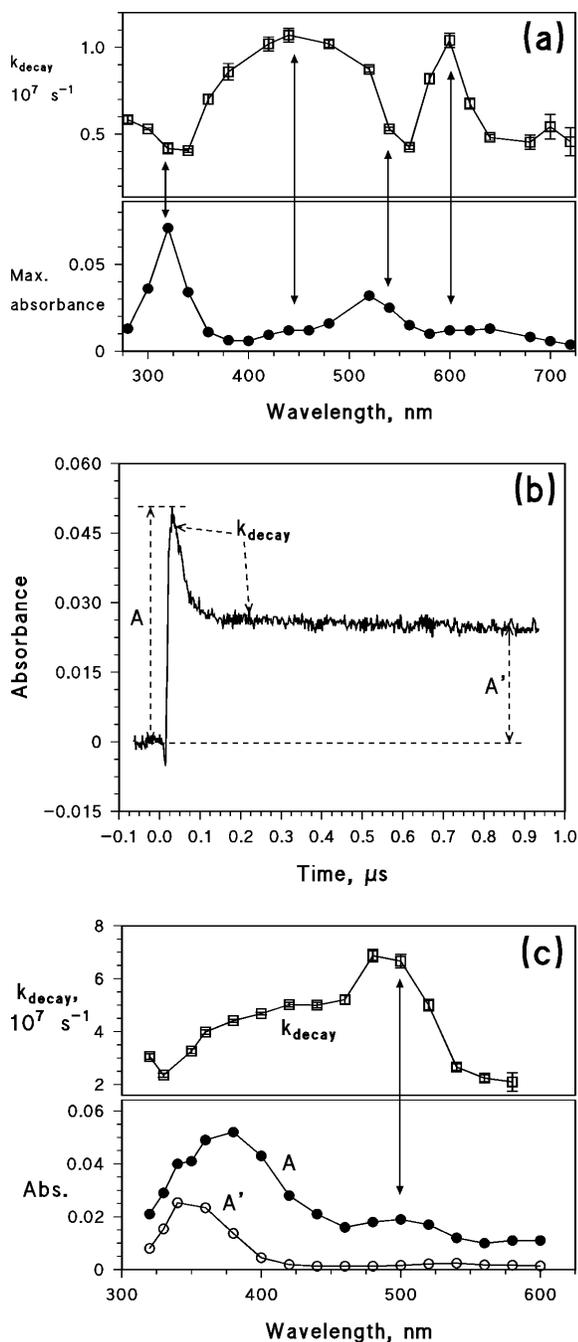
The triplet states of benzophenones often behave as effective photosensitisers for the transformation of phenolic compounds.<sup>46,47</sup> The reactivity between  $^3\text{4BPOH}^*$  and phenol was assessed by studying the effect of phenol concentration on the value of  $k_{520\text{ nm}}$  in acetonitrile. The value of  $k_{520\text{ nm}}$  increases linearly with increasing phenol (see Figure ES14), from which trend a second-order rate constant  $k_{^3\text{4BPOH}^*,\text{Phenol}} = (6.6 \pm 0.3) \cdot 10^7\text{ M}^{-1}\text{ s}^{-1}$  can be obtained between  $^3\text{4BPOH}^*$  and phenol. The reaction yields a species with absorption maximum around 380 nm, which is longer-lived compared to  $^3\text{4BPOH}^*$ . It could be assigned to the phenoxy radical,<sup>48</sup> which suggests that  $^3\text{4BPOH}^*$  would react with phenol by hydrogen abstraction (or by electron transfer followed by deprotonation, which does however look unlikely in aprotic acetonitrile). The formation of phenoxy radicals has indeed been reported upon reaction between carbonyl triplets and phenols.<sup>7</sup>

Additional experiments of laser irradiation showed that negligible reaction would take place between  $^3\text{4BPOH}^*$  and furfuryl alcohol (FFA). Considering that the latter is a  $^1\text{O}_2$  probe,<sup>24,25,38</sup> it would be possible to use FFA to assess the formation of  $^1\text{O}_2$  from irradiated 4BPOH (*vide infra*), without the potential bias of a reaction between FFA and  $^3\text{4BPOH}^*$ .

As additional solvent, 2-propanol was chosen because it has intermediate proticity between aprotic acetonitrile and water.<sup>8,49,50</sup> The initial absorbance of the triplet state after the laser pulse decreased linearly with increasing percentage of 2-propanol (see Figure ES15), suggesting that a lesser amount of  $^3\text{4BPOH}^*$  would be formed in the presence of the alcohol. This result is consistent with a deprotonation of the singlet state in the presence of 2-propanol.

Moreover, the first-order rate constant of  $^3\text{4BPOH}^*$  decay ( $k_{520\text{ nm}}$ ) increased linearly with increasing 2-propanol (see Figure ES16), coherently with a reaction between  $^3\text{4BPOH}^*$  and the alcohol. The corresponding second-order rate constant, derived as the slope of the plot  $k_{520\text{ nm}}$  vs. 2-propanol, is  $k_{^3\text{4BPOH}^*,2\text{-Propanol}} = (3.4 \pm 0.3) \cdot 10^6\text{ M}^{-1}\text{ s}^{-1}$ . The relevant process could be an acid-base and/or a redox reaction, and the literature shows an important disagreement over this issue. While some authors assume that  $^3\text{4BPOH}^*$  is able to abstract hydrogen from alcohols including 2-propanol,<sup>42</sup> others explicitly exclude this possibility and only consider an acid-base process where the alcohol acts as  $\text{H}^+$  acceptor for the deprotonation of  $^3\text{4BPOH}^*$ .<sup>51</sup> For instance, no hydrogen abstraction is reported to take place between  $^3\text{4BPOH}^*$  and ethanol in ethanol solution, where the chemical reactivity of  $^3\text{4BPOH}^*$  should be decreased by its  $\pi\text{-}\pi^*$  configuration.<sup>8</sup>

To get insight into the process details, the reaction with 2-propanol was also studied by laser irradiation of benzophenone (BP), which differs from 4BPOH for the absence of the OH group on the aromatic ring. In this case, it is well known that the BP triplet state abstracts a H atom from the alcohol to form the ketyl radical of BP,<sup>52</sup> which has a comparable lifetime as the BP triplet state and partially overlaps with its absorption spectrum.<sup>53</sup> The laser irradiation of BP was carried out in a system containing 50% acetonitrile and 50% 2-propanol. Figure 5a reports, as a function of the wavelength, the first-order decay constant of the flash photolysis traces ( $k_{\text{decay}}$ , upper graph), as well as the maximum absorbance value reached by each trace, soon after the laser pulse (lower graph). The maximum trace absorbance gives insight into the absorption spectrum of the transient(s) formed by laser irradiation. The absorbance peaks around 320 and 520 nm can be assigned to the triplet state of BP;<sup>43,53</sup> interestingly, these peaks correspond to values of  $k_{\text{decay}} \sim 4 \cdot 10^7\text{ s}^{-1}$ . However, the decay constant shows different values in other wavelength intervals, and in particular it is  $k_{\text{decay}} \sim 1 \cdot 10^8\text{ s}^{-1}$  at around 450 and 600 nm. The variations of  $k_{\text{decay}}$  suggest that the reported maximum trace absorbance is the result of the contribution of more than one species. Literature data indicate that the transient absorption upon laser irradiation of BP and 2-propanol results from both the triplet state and the ketyl radical of BP.<sup>53</sup> Because the BP triplet state absorbs radiation around 320 and 520 nm,<sup>43</sup> one can identify the species with  $k_{\text{decay}} \sim 4 \cdot 10^7\text{ s}^{-1}$  as the triplet state of BP and that with  $k_{\text{decay}} \sim 1 \cdot 10^8\text{ s}^{-1}$  as the ketyl radical. The latter would thus absorb radiation at around 450 and 600 nm. Interestingly, the ketyl radical is not formed upon laser irradiation of BP in pure acetonitrile.



**Figure 5.** (a) Upper plot: first-order decay constant ( $k_{\text{decay}}$ ) of the laser trace as a function of wavelength, upon irradiation (355 nm, 94 mJ/pulse) of 0.1 mM benzophenone in 1:1 2-propanol:acetonitrile. Lower plot: maximum absorbance of the laser trace under the same conditions.

(b) Laser trace (350 nm) upon irradiation of 0.1 mM 4BPOH in 1:1 2-propanol:acetonitrile (355 nm, 94 mJ/pulse). The absorbance values A and A' and the time interval used to determine  $k_{\text{decay}}$  are also highlighted.

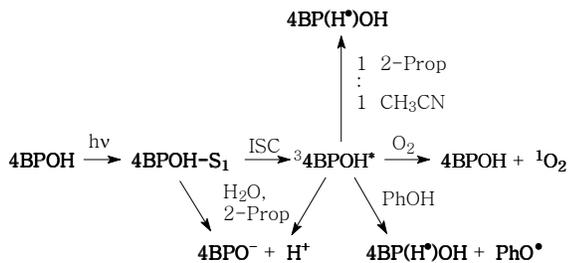
(c) Upper plot: first-order decay constant ( $k_{\text{decay}}$ ) of the laser trace as a function of wavelength, upon irradiation (355 nm, 94 mJ/pulse) of 0.1 mM 4BPOH in 1:1 2-propanol:acetonitrile. Lower plot: wavelength trends of A and A' under the same conditions.

The above-discussed laser irradiation of BP in 50% acetonitrile + 50% 2-propanol yielded transient traces that, after an exponential decay, went back to the absorbance value observed before the laser pulse. The situation changed considerably upon laser irradiation of 4BPOH in the same solvent mixture. In this case, the trace reached a maximum (A) and then decayed exponentially down to a constant value (A') that was higher than the initial absorbance (see Figure 5b). This issue means that a longer-lived species was formed in the presence of 4BPOH, while the exponential decay itself is silent as to the occurrence of one or more short-lived transients.

The plot reported in Figure 5c was obtained under similar conditions as that of Figure 5a, the only difference being that it is referred to 4BPOH instead of BP. In this case as well, the solvent was a 1:1 mixture of acetonitrile and 2-propanol. Also in this case there is evidence of the occurrence of more than one species. First of all, the transient absorption spectrum (A) shows two maxima (around 380 and 500 nm) that, differently from the case of  ${}^3\text{4BPOH}^*$  alone, have quite different absorbance. Secondly,  $k_{\text{decay}}$  varies from  $\sim 2 \cdot 10^7$  to  $\sim 7 \cdot 10^7 \text{ s}^{-1}$ . The plateau absorbance A' has a maximum around 350 nm, which agrees with the absorption spectrum of  ${}^4\text{BPO}^-$ . The latter species is reported to be formed in alcoholic solvents such as ethanol.<sup>8</sup> Further evidence in favour of the identification of the longer-lived species with  ${}^4\text{BPO}^-$  is its decay kinetics, which is very slow in 2-propanol and is accelerated upon addition of water traces (see Figure ESI7). It is reasonable that  ${}^4\text{BPO}^-$ , once formed, undergoes protonation that would be much faster in water compared to 2-propanol.

Figure 5c suggests that, in addition to  ${}^4\text{BPO}^-$  which is relatively long-lived, two transient species with shorter and similar lifetimes occur in the system. The triplet state  ${}^3\text{4BPOH}^*$  is known to absorb at  $\sim 330$  and  $\sim 500\text{--}520 \text{ nm}$ ,<sup>43</sup> thus it should be the species with  $k_{\text{decay}} \sim 7 \cdot 10^7 \text{ s}^{-1}$ . The other species would account for the observed  $k_{\text{decay}} \sim 2 \cdot 10^7 \text{ s}^{-1}$  at around 575 nm, and possibly also around 320 nm. In analogy with the results obtained with BP, the second species might be the ketyl radical of 4BPOH ( ${}^4\text{BP}(\text{H}^*)\text{OH}$ ). This radical is reported to have an absorption maximum at 560 nm,<sup>6</sup> which would be reasonably consistent with the observed wavelength trend of  $k_{\text{decay}}$ . If the identification of the second species with  ${}^4\text{BP}(\text{H}^*)\text{OH}$  is correct, there would be support for the possibility of an electron-transfer reaction between  ${}^3\text{4BPOH}^*$  and 2-propanol. However, one should also consider that the experiments reported in Figure 5c were carried out in 1:1 2-propanol:acetonitrile. In pure acetonitrile there is an equilibrium between the  $n\text{-}\pi^*$  and  $\pi\text{-}\pi^*$  configurations of  ${}^3\text{4BPOH}^*$ ,<sup>6</sup> which allows for the occurrence of the photochemically reactive  $n\text{-}\pi^*$ . The equilibrium could be maintained in mixtures of acetonitrile and 2-propanol, while in pure 2-propanol one cannot exclude the occurrence of  $\pi\text{-}\pi^*$  alone. Unfortunately, flash photolysis experiments in pure 2-propanol are inconclusive because the lack of  ${}^3\text{4BPOH}^*$  detection could be due either to a very short lifetime of  $\pi\text{-}\pi^*$ , or (more likely) to a very fast deprotonation of  ${}^3\text{4BPOH}^*$  (in either configuration) to give  ${}^4\text{BPO}^-$ .

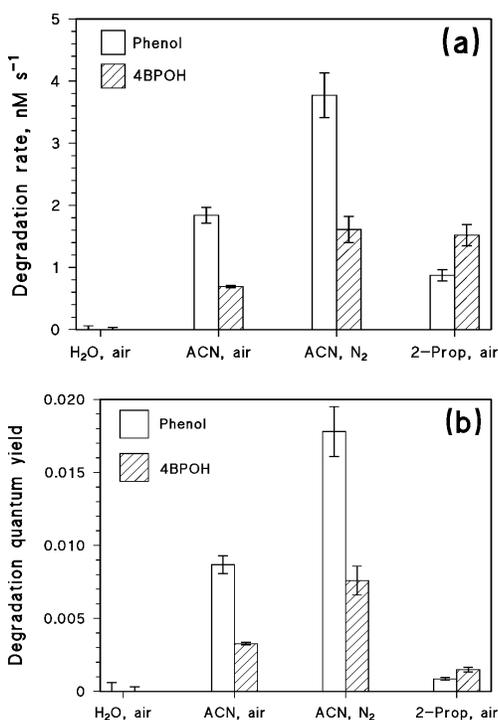
Scheme 1 shows the hypothesised processes involving irradiated 4BPOH. Excited-state deprotonation and  $4\text{BPO}^-$  protonation would be faster in water than in 2-propanol.



**Scheme 1.** Hypothesised processes involving the transient species formed upon irradiation of 4BPOH. 2-Prop: 2-propanol; PhOH: phenol;  $\text{PhO}^\bullet$ : phenoxy; ISC: inter-system crossing.

### Steady irradiation experiments

The ability of 4BPOH to photosensitise the degradation of phenol was tested in water, acetonitrile and 2-propanol. The transformation rates  $R_s$  of the two substrates ( $S = 4\text{BPOH}$  or phenol, both at 0.1 mM initial concentration) under the different conditions are reported in Figure 6a. Figure 6b reports the relevant quantum yields of transformation (calculated as  $\Phi_s = R_s (P_a^{4\text{BPOH}})^{-1}$ ).



**Figure 6.** Degradation rates (a) and degradation quantum yields (b) of phenol and 4BPOH in mixture (initial concentration 0.1 mM for both compounds), upon UVB irradiation in different solvents (water, acetonitrile, 2-propanol) and different conditions (air or nitrogen atmosphere). The error bars represent  $\pm\sigma$ . The rates and quantum yields in water were negligible, and only the related uncertainty (detection limit) is reported. ACN = acetonitrile

No degradation of either 4BPOH or phenol was observed in water, while both compounds were photodegraded in the organic solvents. Focusing first on the degradation rate of phenol, in aerated solution it followed the order acetonitrile > 2-propanol > water. An analogous trend was observed for the quantum yields, just with larger differences between acetonitrile and 2-propanol compared to the rates: the reason is that 4BPOH absorbs lamp radiation to a higher extent in 2-propanol compared to acetonitrile ( $P_a^{4\text{BPOH}, 2\text{-propanol}} > P_a^{4\text{BPOH}, \text{CH}_3\text{CN}}$ ). Moreover, phenol degradation was enhanced with acetonitrile under nitrogen atmosphere. These results can be explained easily in the framework of the photochemical processes depicted in Scheme 1, considering that  $^34\text{BPOH}^*$  would oxidise phenol to the phenoxy radical and, therefore, trigger phenol transformation. Additional evidence for the formation of the phenoxy radical is the occurrence of the compound 4PP (formed upon dimerisation of phenoxy<sup>54</sup>) upon irradiation of 1mM phenol + 1 mM 4BPOH in acetonitrile (15-20  $\mu\text{M}$  4PP was produced for irradiation times > 24h) and 2-propanol (1-1.5  $\mu\text{M}$  4PP was formed at the same irradiation time scale).

The lack of phenol degradation in water would be accounted for by the fast deprotonation/quenching of the 4BPOH excited states in this solvent. Indeed, no signal of  $^34\text{BPOH}^*$  could be detected upon laser irradiation of 4BPOH in aqueous solution. In the case of 2-propanol as solvent, the reaction between  $^34\text{BPOH}^*$  and phenol would be in competition with triplet deprotonation and possibly with the oxidation of the solvent (although some evidence of the oxidation of 2-propanol by  $^34\text{BPOH}^*$  is only available for mixtures with acetonitrile, and not in pure alcohol). These reactions would reduce the availability of  $^34\text{BPOH}^*$  for the degradation of phenol in 2-propanol, but they would not be as fast as the decay of excited 4BPOH in water. Finally, because  $^34\text{BPOH}^*$  does not react with acetonitrile, phenol photosensitised degradation was fastest in this solvent.

The faster degradation of phenol in the absence of oxygen (acetonitrile as solvent, Figure 6), due to reaction with  $^34\text{BPOH}^*$ , is accounted for by the quenching of  $^34\text{BPOH}^*$  by oxygen to yield  $^1\text{O}_2$ . Coherently, Figure ESI1 suggests that the decay constant of  $^34\text{BPOH}^*$  in acetonitrile gets about double when passing from a deoxygenated to an aerated system. This is fully consistent with the observed degradation of phenol, which was about twice faster under nitrogen atmosphere compared to air (Figure 6). In the aerated system, 0.1 mM phenol would scavenge approximately 0.06% of  $^34\text{BPOH}^*$ , a percentage that would become approximately double in the absence of oxygen. However, in aerated solution there would be formation of  $^1\text{O}_2$  that could also contribute to phenol degradation. The first-order decay constant of  $^1\text{O}_2$  in acetonitrile is about  $2.5 \cdot 10^4 \text{ s}^{-1}$ ,<sup>55</sup> while the second-order reaction rate constant between phenol and  $^1\text{O}_2$  seems to be strongly solvent-dependent. Indeed, while rate constants of  $10^6$ - $10^7 \text{ M}^{-1} \text{ s}^{-1}$  are reported for water,<sup>56</sup> in organic solvents one has rate constant values of about  $10^4 \text{ M}^{-1} \text{ s}^{-1}$ .<sup>57,58</sup> The difference could at least in part be explained by the higher

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reactivity of the phenolate anion toward  $^1\text{O}_2$ , coherently with the fact that the rate constant in water increases with increasing pH.<sup>54</sup> If the reaction rate constant between phenol and  $^1\text{O}_2$  in acetonitrile were in the range of  $10^4 \text{ M}^{-1} \text{ s}^{-1}$ , only ~0.004% of  $^1\text{O}_2$  would react with 0.1 mM phenol and the reaction with singlet oxygen would play a secondary role in phenol degradation (~7% of the total). Triplet 4BPOH would thus induce more effective degradation of phenol under  $\text{N}_2$  atmosphere, where triplet quenching by  $\text{O}_2$  would not be operational.

Focusing now on the degradation of 4BPOH in the different solvents, one observes first of all the absence of transformation in water, which can be explained as per the above discussion. The 4BPOH degradation in acetonitrile was faster in the absence of oxygen, which quenches  $^3\text{4BPOH}^*$  to give back the ground-state molecule in a null cycle of excitation-deactivation:



The degradation rate of 4BPOH was higher in 2-propanol than in acetonitrile (Figure 6a). One should consider that  $p_a^{4\text{BPOH}, 2\text{-Propanol}} > p_a^{4\text{BPOH}, \text{CH}_3\text{CN}}$  (by a factor of ~4.5) and, as far as the transformation quantum yields are concerned, it was  $\Phi_{4\text{BPOH}}^{\text{CH}_3\text{CN}} > \Phi_{4\text{BPOH}}^{2\text{-Propanol}}$  (Figure 6b). Therefore, while the difference in the rates can be accounted for by the higher radiation absorption of 4BPOH in 2-propanol, the lower quantum yield in the alcoholic solvent is consistent with the quenching of  $^3\text{4BPOH}^*$  by an acid-base and possibly also a redox process.

Finally, the formation of  $^1\text{O}_2$  by irradiated 4BPOH (0.1 mM initial concentration) was assessed in water, acetonitrile and 2-propanol by addition of 0.1 mM FFA. The initial rate of FFA transformation ( $R_{\text{FFA}}$ ) was negligible in water, and it was  $R_{\text{FFA}} = (8.66 \pm 0.68) \cdot 10^{-10} \text{ M s}^{-1}$  in acetonitrile and  $R_{\text{FFA}} = (1.11 \pm 0.14) \cdot 10^{-10} \text{ M s}^{-1}$  in 2-propanol. Because  $^3\text{4BPOH}^*$  shows negligible reactivity with FFA, as suggested by LFP experiments, the observed degradation of FFA would be accounted for by the production of  $^1\text{O}_2$  in the irradiated systems.

## Conclusions

The singlet and triplet excited states of 4BPOH undergo deprotonation in the presence of water to produce ground-state  $4\text{BPO}^-$ , which quickly adds  $\text{H}^+$  to yield back the ground-state 4BPOH. This process quenches the fluorescence and the photosensitising activity of 4BPOH in aqueous solution. In aprotic acetonitrile, 4BPOH shows significant fluorescence emission with a fluorescence band (Ex/Em ~ 325/475 nm) that overlaps with peak C of humic substances. In acetonitrile,  $^3\text{4BPOH}^*$  mainly reacts with  $\text{O}_2$  to produce  $^1\text{O}_2$  and it is also involved in the degradation of phenol, with a second-order reaction rate constant  $k_{^3\text{4BPOH}^*, \text{Phenol}} = (6.6 \pm 0.3) \cdot 10^7 \text{ M}^{-1} \text{ s}^{-1}$ . The reaction would produce the phenoxy radical, and presumably also the ketyl radical of 4BPOH.

With the solvent 2-propanol, which shows some protic properties and can be oxidised, light-excited 4BPOH undergoes

partial deprotonation to  $4\text{BPO}^-$  (but to a considerably lesser extent than in water), and very slow reprotonation. There is also some evidence that  $^3\text{4BPOH}^*$  could abstract an H atom from 2-propanol in alcohol:acetonitrile mixtures. In pure 2-propanol, it is possible that  $^3\text{4BPOH}^*$  is in a  $\pi\text{-}\pi^*$  configuration that is non-reactive toward the oxidation of the alcohols. However, in pure 2-propanol one does not observe  $^3\text{4BPOH}^*$  signals at least because of the fast deprotonation, thus this work is silent about the ability of  $^3\text{4BPOH}^*$  to oxidise 2-propanol in pure alcoholic solvent. The partial  $^3\text{4BPOH}^*$  quenching in 2-propanol, whatever the actual pathway(s), decreases the photosensitising ability of 4BPOH in 2-propanol compared to acetonitrile. However, the ability of irradiated 4BPOH to produce  $^1\text{O}_2$  and degrade phenol is still significant in the alcoholic solvent.

Considering that benzophenones (possibly including 4BPOH) are important constituents of CDOM and particularly of humic substances, the present findings may be significant for the photoactivity of the hydrophobic and hydrophilic humic moieties. Because a compound such as 4BPOH would not be fluorescent and would not produce  $^1\text{O}_2$  in the presence of water, but would do so in its absence, it might contribute to the occurrence of elevated  $^1\text{O}_2$  levels in the hydrophobic cores of humic material.

## Acknowledgements

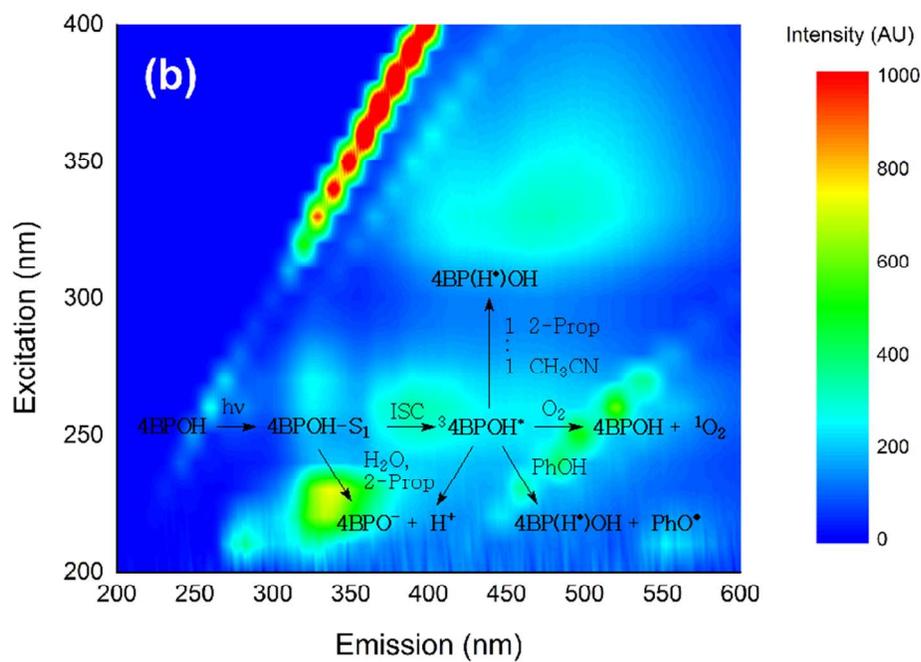
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