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## Operando XAFS of supported Pd nanoparticles in flowing ethanol/water mixtures: implications for catalysis.

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**Ethanol-water, a prototypical 'green' solvent mixture, cannot be considered as inert toward supported Pd nanoparticles. We establish severe size dependencies in Pd redox behavior, persistent gradients in Pd phase in a plug-flow situation, and that Pd nanoparticles are contaminated with hydrogen derived from the solvent.**

Atom efficient catalysis and continuous flow methods are important enabling technologies for the manufacture of fine chemicals and pharmaceuticals, where the use of solvent is ubiquitous and counts as the single greatest contributor to overall environmental impact of these processes.<sup>1-6</sup> The choice of a solvent is primarily driven by physicochemical properties (polarity, protic or aprotic) that are compatible with the reactants, while rendering the appropriate environment required to deliver the optimal rate and selectivity. In recent years, there have been several monographs<sup>6-11</sup> produced by the pharmaceutical industry to assess and clarify the use of organic solvents. In all of these reviews, water and ethanol consistently scored very favourably by existing metrics.

Perhaps as a result of this, ethanol-water is used extensively in a myriad of Pd-catalyzed reactions, including redox processes<sup>12</sup> and all of the major cross-coupling reactions.<sup>13-16</sup> As such we have studied how supported palladium – starting from nanoparticulate PdO – behaves when exposed to mixtures of ethanol and water. Through time and spatially resolved Extended X-ray Absorption Fine Structure (EXAFS) we attempt to quantify and understand fundamental solvent-Pd interactions and how they influence the development and behaviour of the Pd as a function of its loading within the paradigm of a continuous plug flow reactor.

EXAFS provides a powerful means to directly address the oxidation states and structure of Pd *in operando* with high temporal and spatial resolution. There now exist a number of examples of this technique being applied to study various aspects of Pd catalyzed selective oxidation<sup>17-22</sup> and homogeneous<sup>23-28</sup> and heterogeneous<sup>29, 30</sup> catalytic carbon-

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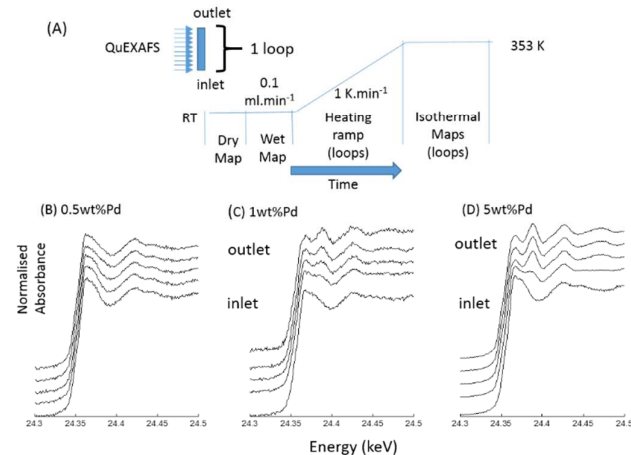
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carbon coupling. However, these operando EXAFS studies have traditionally employed recirculating batch reactors. Here, our studies are performed using a continuous flow reactor,<sup>30</sup> which is industrially more relevant in terms of scale up and ease of operation; for this reason alone research into flow based processing solutions has expanded rapidly in recently years.<sup>31–33</sup> Within the context of the present study, a continuous flow arrangement can also provide both spatial and temporal windows into fundamental chemistry occurring within supported Pd system, making them easier to observe and quantify compared to a batch situation.

Using this method we demonstrate that when in contact with a mixture of ethanol and water Pd exhibits very large loading (particle size) dependencies in its redox behaviour that lead to persistent axial gradients in the structure of the metal phase. Moreover, even when metallic Pd is formed, the Pd nanoparticles are significantly contaminated with stored hydrogen. This hydrogen arises from reactive chemistry occurring between the catalyst and the solvent mixture. The desired situation, wherein the solvent for a catalytic process is an inert carrier of other reactive species, is never obtained in this prototypical case. The ramifications of these observations for the continued “greening” of such chemistry using supported Pd nanoparticles are discussed.

Figure 1 (A) shows the overall experimental approach with figures 1(B–D) showing examples of normalised Quick EXAFS (QuEXAFS) spectra obtained at the end of the experiment for supported Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> catalysts of 0.5, 1, and 5wt% Pd loading.



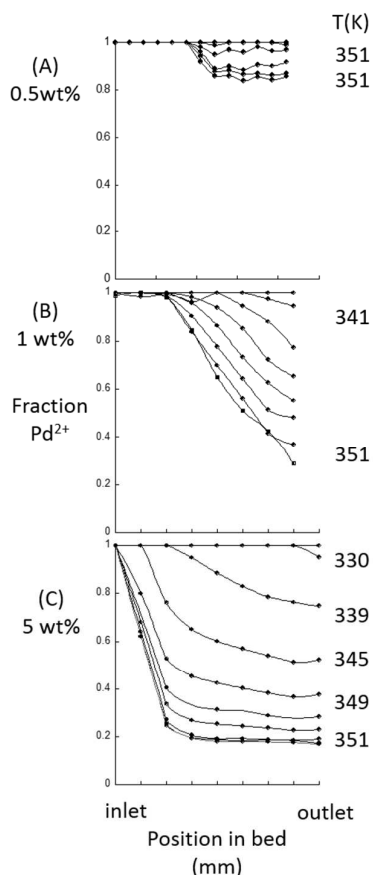
**Figure 1.** (A) Experimental method (see also SI). Transmission QuEXAFS spectra are repeatedly collected at different axial positions along the catalyst bed whilst it is heated to, and then maintained at, 353 K under a flow (0.1 ml·min<sup>-1</sup>) of 50:50 EtOH/H<sub>2</sub>O. (B–D) Representative and normalised Pd K-edge XAFS from different axial positions within catalyst beds comprised of (B) 0.5wt%Pd/Al<sub>2</sub>O<sub>3</sub>; (C) 1wt%Pd/Al<sub>2</sub>O<sub>3</sub>; and, (D) 5wt%Pd/Al<sub>2</sub>O<sub>3</sub>.

Briefly, having established a flow of solvent through the catalyst bed transmission Pd K edge Quick EXAFS measurements are made sequentially along the length of the bed to create maps of the chemical state and structure of the Pd in the catalyst. This mapping process is continued as the sample is then heated to a maximal temperature of 353K and continued for some time at this temperature (see also SI).

In their native form, all of the Pd catalysts contains significant amount of oxidised PdO, which undergoes reduction to Pd(0) by ethanol. What is interesting, and immediately evident from figure 1, is that neither the chemical state of the Pd in these systems, nor the manner in which it changes axially along the catalyst bed, is the same when the Pd loading is changed. The same experiment conducted using H<sub>2</sub>O alone elicits no reduction of the PdO in any of the cases shown. As such the observed behaviour is ascribed to chemistry occurring ostensibly between the ethanol and the supported Pd.

By the end of our experiment hardly any of the Pd in the 0.5wt% case appears to have changed from its native oxidised state. In stark contrast, at 5wt%, Pd practically all of the catalyst bed has reduced to yield metallic Pd with oxidised Pd existing only at the reactor inlet. The 1wt% case sits somewhere in between evidencing a more gradual shift from highly oxidised to predominantly reduced.

Figure 2 (a–c) shows, for the same three systems, results of a linear combination analysis (LCA) of the Pd K edge XANES to assess the fractions of oxidised and reduced Pd present as a function of time, temperature, and position within the bed.



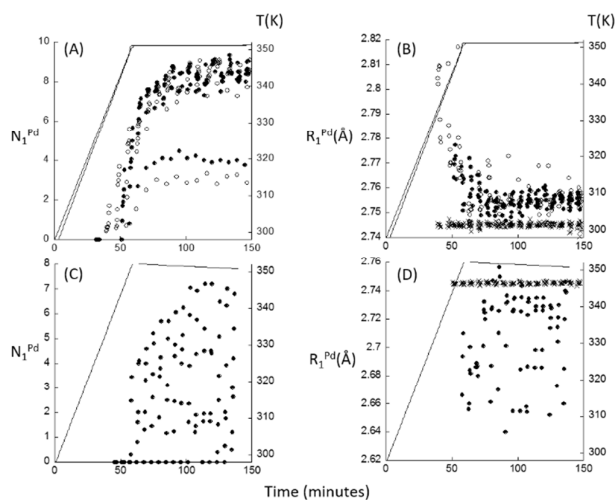
**Figure 2.** Results of LCA analysis of the Pd K edge XANES for the approximate fraction of Pd in the Pd<sup>2+</sup> state as a function of time and temperature for the 0.5, 1, and 5wt% Pd/Al<sub>2</sub>O<sub>3</sub> cases shown in figure 1. The references used for the analysis were a Pt foil and bulk PdO powder.

This analysis must be regarded as approximate as the references used are bulk metal foils and PdO. Nevertheless, it

clearly shows the magnitude of the loading/size dependencies in the reduction process, and that a uniformly reduced catalyst bed is never attained. It also delineates from a spatial perspective how the reduction takes place, i.e. in the opposite direction of the flow from the reactor outlet to the inlet. This is commensurate with an earlier observation made during gas phase methane oxidation over Pt/Al<sub>2</sub>O<sub>3</sub> catalysts at high temperature (603K),<sup>33</sup> and strongly indicative of cumulative reduction potential arising as the solvent passes along the bed in the direction of flow.

From the above it is clear that heating to 353 K under flowing EtOH/H<sub>2</sub>O (50:50) can elicit the reduction of PdO to metallic Pd. It is also apparent that this is highly dependent on the average size of the Pd nanoparticles (see also SI) and that, within an environment under continuous flow of an alcoholic solvent, persistent Pd-phase gradients develop backwards from the reactor exit toward the inlet.

Figure 3 shows the results of analysing the EXAFS obtained during these experiments for 1, 5, and 10 wt% systems. Figure 3(A) reports the first shell coordination numbers pertaining to any fcc, metallic Pd phase ( $N_{1}^{Pd}$ ) as a function of time/temperature, for 10 wt% (open circles) and 5wt% systems (filled circles), whilst figure 3(B) reports the first shell PdPd bondlengths ( $R_{1}^{Pd}$ ) returned for these samples along with those derived for a Pd foil (crosses) used as an internal reference. Figures 3 (C) and (D) show the same data for a 1



wt% Pd system.

**Figure 3.** (A) and (B) Comparison of the behaviour of 5 (filled circles) and 10 (open circles) wt% Pd/Al<sub>2</sub>O<sub>3</sub> catalysts during heating under flowing EtOH/H<sub>2</sub>O (50:50) from analysis of the Pd K-edge EXAFS. (A) 1<sup>st</sup> shell, fcc Pd coordination number ( $N_{1}^{Pd}$ ) and (B) 1<sup>st</sup> shell PdPd bond length ( $R_{1}^{Pd}$ ). In each panel the solid black lines show the heating ramps (right hand abscissa). Panels (C) and (D) report the same parameters but for a 1wt% Pd/Al<sub>2</sub>O<sub>3</sub> sample. In panels (B) and (D) the crosses are the  $R_{1}^{Pd}$  values returned from analysis of the reference Pd foil (B). These plots contains all data obtained for each loop and each axial position, from inlet to outlet, in the bed and therefore also reflect the axial gradients in structure of the Pd phase as they develop.

EXAFS analysis yields more information as to the development of the reduced Pd phase. Firstly, the onset of PdO reduction, and formation of fcc Pd nanoparticles, occur at lowest temperature in the 10% case and the highest in the 1wt%.

Secondly, the inlet of the bed remains to varying degrees oxidised in each case, with the 1wt% showing no evidence of reduced Pd at the reactor inlet throughout the experiment. Figure 3 also shows that the 1wt% case, as might be expected, eventually yields Pd particles that are very much smaller than observed in the 5 and 10 wt% cases.<sup>35</sup> The latter pair, within the error associated with this measurement, appear very similar in terms of the size of the average Pd nanoparticle implied to be formed by the end of the experiment.

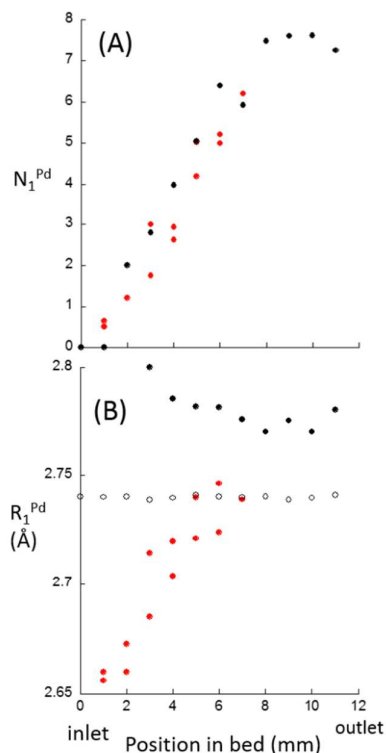
When reduced 5 and 10 wt% Pd catalysts show values of  $R_{1}^{Pd}$  that are consistently larger than those obtained from the Pd reference foils. Previous work<sup>36</sup> has established that clean Pd nanoparticles of the size implied from the values of  $N_{1}^{Pd}$  (and TEM, see SI) should never show PdPd bond lengths longer than found in a foil. In the 1wt% case, when maximally reduced toward the outlet of the reactor, average values of ca. 2.73–2.74 Å are indicated. Although these might also be considered large for the size of the Pd particles within this sample, it is much less easy to establish whether they are equally anomalous.

In the case of Pd,<sup>36</sup> and in the absence of other factors, smaller particle sizes are accompanied by smaller bond distances. As such, the expanded bond lengths in the 5 and 10 wt% cases result from the process of reduction of the Pd by the solvent. There are two possible explanations for this: first, adsorbates may form and bind to the surfaces of the nanoparticles lengthening the surface PdPd bonds to such a degree that the average bonding within the particle, as observed by the EXAFS, appears extended; or, the Pd nanoparticles may have expanded as a result of incorporation of another element. The ability for Pd to absorb atomic hydrogen (to form Pd H<sub>x</sub> phases) or carbon (to yield PdC<sub>x</sub>) is well established,<sup>37–43</sup> and both of these interstitial compounds are known to be thermally stable at the temperatures we have applied.

Figure 4 shows the results obtained from a further experiment. In this case (black) the 1 wt% Pd system was maintained at 353 K under flowing EtOH/H<sub>2</sub>O for 11 hours before the system was exposed to X-rays for mapping of the catalyst bed at room temperature. The results of EXAFS analysis of this mapped bed are compared to those derived from the very end (last two loops) of the experiment already described (figures 1 and 2, ca. 1.5 hours at 353 K in the X-ray beam- red). Figure 4(A) reports the axial variation in  $N_{1}^{Pd}$  in the two cases, figure 4(B) the corresponding variations in  $R_{1}^{Pd}$ . Again the result of the fitting the Pd reference foil are also given (open circles) in Figure 4(B).

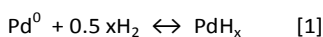
Figure 4(A) reveals that the results for  $N_{1}^{Pd}$  are practically identical. In the 1wt% case, the evolution of the catalyst bed appears to become stuck; the axial gradients in reduced Pd seen after 1.5 hours (see figure 2) remain the same even after a further 10 hours of operation. The values of  $R_{1}^{Pd}$  (figure 4(B)) obtained after 1.5 or 11 hours of flow operation show a marked difference. After 11 hours the PdPd bond length has increased significantly (ranging from 2.8 to 2.77 Å) and the axial gradient in  $R_{1}^{Pd}$  has flattened.

These experiments demonstrate that the chemistry being observed is not due to the X-rays themselves. Furthermore, these latter observations establish the extended PdPd bonding as the result of a storage of atomic hydrogen; it is only by the formation of such PdH<sub>x</sub> phases that such extended PdPd bonding within can be reasonably explained.



**Figure 4.** Comparison of  $N_1^{Pd}$  (A) and  $R_1^{Pd}$  (B) as a function of axial position along the reactor bed for 1wt%Pd/Al<sub>2</sub>O<sub>3</sub> heated under flowing EtOH/H<sub>2</sub>O and maintained at 353 K under the flowing solvent for ca. 1.5 hours (red) and 11 hours (black). In figure 3(B) the open circles report  $R_1^{Pd}$  obtained from the reference Pd foil. The sample run for 1.5 hours was under constant X-ray illumination. The 11 hour experiment was run in the absence of X-rays and only exposed post reaction for the duration of the mapping process (ca. 15 minutes) at 298 K.

The differences in  $R_1^{Pd}$  values shown in figure 4(B) and the time/temperature dependence of the  $R_1^{Pd}$  curves for the 5 and 10% Pd systems (figure 3) can now also be understood as reflecting the thermal stability of the PdH<sub>x</sub> phase and the establishment of the following equilibrium in these systems from the very earliest stages of Pd reduction:

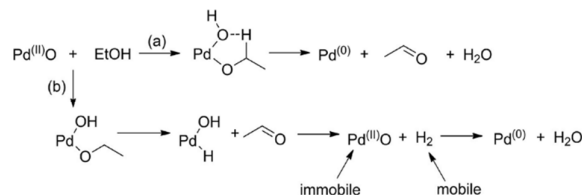


In the 5 and 10 wt% cases it appears that the Pd nanoparticles initially formed are heavily contaminated with hydrogen. Over similar Pd/Al<sub>2</sub>O<sub>3</sub> and Pd/C catalysts, using EXAFS as a probe, McCauley<sup>38</sup> established that a PdH<sub>x</sub> (x=0.44) phase has  $R_1^{Pd}$  bond lengths of 2.81–2.82 Å, which (from Figure 3) are also attained in the very earliest stages of the reduction of the 10wt% Pd sample. As the temperature of the sample is increased equilibrium [1] gradually shifts to the left, as demonstrated by figure 3, but never to the point where the

reduced Pd nanoparticles can be considered free from hydrogen contamination.

Our experiments therefore reveal and quantify two separate consequences of exposing supported Pd, starting with nanoparticulate PdO, to a prototypical green solvent mixture. The results reveal large differences in redox behaviour over the range of Pd content investigated. For a 0.5 wt% Pd catalyst (figures 1 and 2), no evidence was found for the formation of a discrete metallic phase. This contrasts with the relatively facile reduction of Pd that occurs at loadings of ≥5wt%. However, even at these loadings, a small proportion of the bed at the reactor inlet remains significantly oxidised. A homogeneous axial Pd phase distribution is therefore never achieved and, in the limit of high Pd dispersion, severe gradients in oxidic/metallic Pd arise that are extremely persistent.

We envisage that the initial reduction of PdO by ethanol can occur primarily in two ways as shown in scheme 1; we have previously suggested<sup>30</sup> that the second route (b) is more likely, as it could result in a cumulative reduction potential along the bed in the direction of the flow and thus lead to the types of steady state gradients observed.



**Scheme 1.** Two possible routes for ethanol to reduce PdO and form Pd nanoparticles which may further oxidise solvent molecules and/or store and release H subsequently eliciting further Pd reduction downstream.

Analysis of the EXAFS (rather than simple mapping from XANES<sup>30</sup>) shows that this scheme needs to be augmented to take account of an initial high level of hydrogen storage. In addition Pd nanoparticles will themselves be able to catalyse reduction of ethanol further feeding further feed into equilibrium [1] to accelerate downstream reduction of more Pd.

Investigations with a 2 wt% catalyst (see SI) show its behaviour is more akin to that displayed by the 1 wt% Pd case than the 5 wt%. This suggests that the redox properties of supported Pd display a step change as a function of particle size (between ca. 3–4 nm, see SI). This change in behaviour is the foundation for the observed differences between these catalysts.

Evidence for size-dependent changes in the properties of Pd has been previously obtained from ex-situ XPS measurements of Pd/C catalysts used in aerobic oxidation of glucose.<sup>44</sup> The redox behaviour of Pd nanoparticles in high temperature CO oxidation<sup>45</sup> has also been observed to change abruptly between a Pd loading of 4wt% (ca. 4 nm average particle diameter) and 2 wt% (ca. 3nm average particle diameter). Given the very different experimental conditions involved in these and the current case, it would seem that switching of the chemical properties of Pd nanoparticles occurs somewhere between ca. 3 and 4 nm, which has important implications for their catalytic applications. The fundamental behaviour we have observed of supported Pd interacting with an EtOH/H<sub>2</sub>O



solvent also has further ramifications for the adoption of this system on a large scale. Firstly, the induction of phase gradients under flow must be considered as deleterious to any applied process situation. Secondly, reducing the loading of Pd may not result in the desired atom efficiency one might expect. Our results show that below ca. 3nm particle size the supported Pd will have a strongly enhanced tendency to remain oxidised which may or may not be appropriate for the desired process chemistry.

Further, the contamination of Pd nanoparticles by hydrogen may have a number of consequences: theoretical studies have shown that subsurface H in Pd nanoparticles may significantly affect the stability of adsorbates on the surfaces on these particles;<sup>46</sup> particle morphologies, and therefore the number and type of potential reactive sites, themselves may also be altered as a result of such bulk contamination. As such the formation of PdH<sub>x</sub>, rather than clean Pd, may have profound effects in directing the overall catalytic properties of the system.

Equilibrium [1] demonstrates that PdH<sub>x</sub> should be removable from such systems through operation at higher temperatures. However, this may also promote undesirable leaching of Pd.<sup>29</sup> Within the development of green processes this solution may also be regarded as self-defeating as it requires a greater energy input and thus lower overall process efficiency.

In summary, the use of a prototypical green solvent mixture of a primary alcohol and water has a number of reactive ramifications for supported Pd nanoparticles that must be understood before one may start to consider the suitability of such systems for catalysis. These are: extreme changes in the basic redox behaviour of the Pd as a function of particle size resulting in persistent and undesirable gradients in Pd structure within a flow reactor; and, extensive formation of PdH<sub>x</sub>. Whilst some of these factors could be avoided through higher temperature and pressure operation this may not be desirable for the overall process efficiency. At least in the case of supported Pd, it seems clear that more basic research into understand how such differing factors may be controlled may provide the best long term route for the attainment of greener process solutions for fine chemical synthesis.

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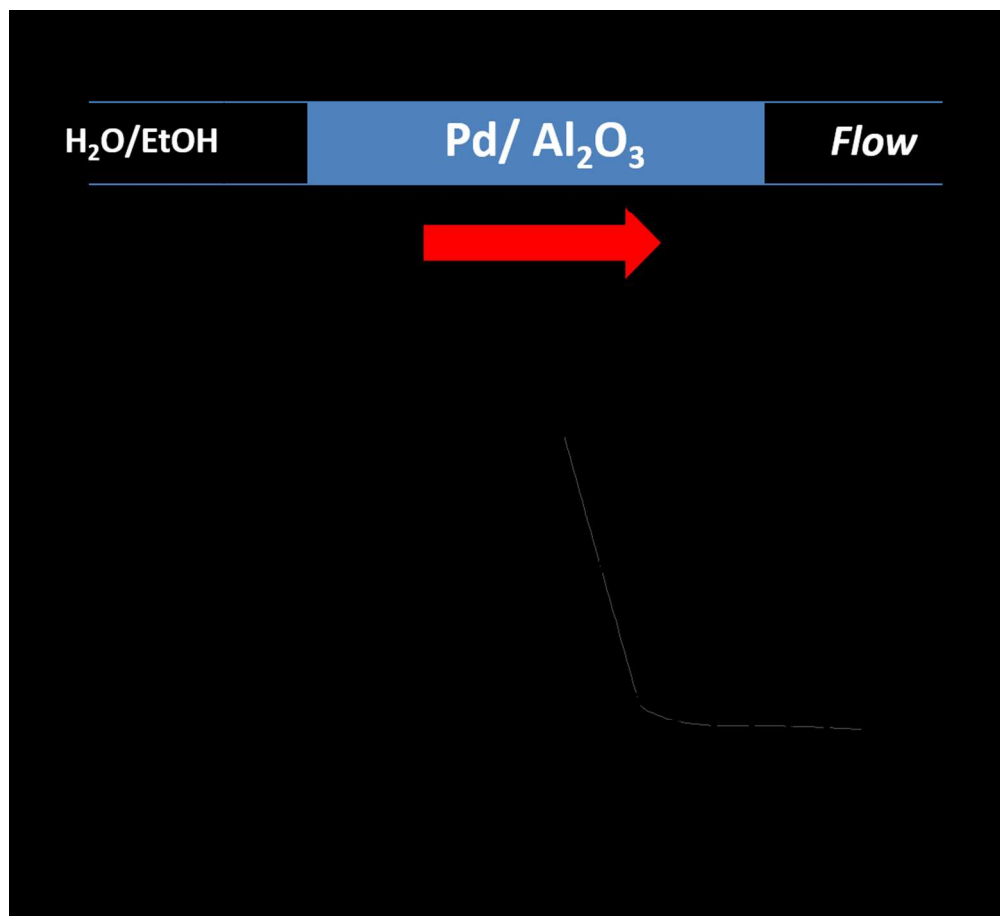
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192x175mm (150 x 150 DPI)