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ARTICLE

## In situ ATR-FTIR study of H<sub>2</sub>O and D<sub>2</sub>O adsorption on TiO<sub>2</sub> under UV irradiation†

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Received 00th January 20xx,  
Accepted 00th January 20xx

DOI: 10.1039/x0xx00000x

www.rsc.org/

The adsorption of water and deuterium oxide on TiO<sub>2</sub> surfaces was investigated in the dark as well as under UV(A) irradiation using in situ ATR-FTIR spectroscopy under oxygen and oxygen free conditions. Adsorption of H<sub>2</sub>O-D<sub>2</sub>O mixtures revealed an isotopic exchange reaction occurring onto the surface of TiO<sub>2</sub> in the dark. Under UV(A) irradiation, the amount of both OH and OD groups was found to be increased by the presence of molecular oxygen. Furthermore, the photocatalytic formation of hydroperoxide under oxygenated condition has been recorded utilizing Attenuated Total Reflection Fourier Transformed Infrared (ATR-FTIR) spectroscopy which appeared as new band at 3483 cm<sup>-1</sup>. Different possible mechanisms are discussed in terms of the source of hydroxyl groups formed and/or hydration water on the TiO<sub>2</sub> surface for the photocatalytic reaction and photoinduced hydrophilicity.

### 1. Introduction

Photocatalytic reactions at the surface of titanium dioxide (TiO<sub>2</sub>) have been attracting much attention, in view of their practical applications including environmental aspects such as water purification, self-cleaning and antifogging.<sup>1,2</sup> UV(A) irradiation of the TiO<sub>2</sub> surface induces photocatalytic reactions and photoinduced hydrophilicity.<sup>3</sup> In general, electrons and holes generated after the band gap excitation can respectively, reduce and oxidize electron acceptor and electron donor species which are already adsorbed on the TiO<sub>2</sub> surface. The importance of the photogenerated active oxygen species, such as hydroxyl radicals (<sup>•</sup>OH), superoxide radicals (O<sub>2</sub><sup>•-</sup>) and hydrogen peroxide (H<sub>2</sub>O<sub>2</sub>) in photocatalytic applications has been extensively studied.<sup>4,5</sup> The mechanism of hydroxyl radical formation has been suggested as a result of the oxidation of water,<sup>6</sup> while hydrogen peroxide can be formed by both reduction and oxidation processes, such as dimerization of hydroxyl radicals, as well as disproportionation and reduction of superoxide radical anions.<sup>5</sup> On the other hand, the photoinduced hydrophilicity is reported as a result of increasing the number of hydroxyl groups on the TiO<sub>2</sub> surface upon irradiation.<sup>7</sup> During these processes, the surface OH groups on the TiO<sub>2</sub> surface have been assumed to play important roles in both photocatalytic reactions and photoinduced hydrophilicity.<sup>8</sup> Adsorbed species such as water and molecular oxygen, however are expected to affect the hydroxyl group

behaviour on the TiO<sub>2</sub> surface,<sup>9</sup> but their influences have to date not yet been clearly determined.

Several mechanisms regarding the interaction of UV(A) light and hydroxyl group formation on TiO<sub>2</sub> have been proposed by different research groups.<sup>8</sup> One of the proposed mechanisms for the photoinduced hydrophilicity is based on the formation of oxygen vacancies and the reconstruction of Ti-OH bonds caused by the photogenerated electron-hole pairs.<sup>7,10</sup> Another mechanism has been suggested to involve the photocatalytic removal of organic contaminants from the surface.<sup>11</sup> Hashimoto and co-workers have followed the idea of photogenerated surface OH groups to introduce their surface reconstruction model of hydroxyl groups as the mechanism for the hydrophilic conversion.<sup>12,13</sup> Another mechanism was proposed by Yates *et al.*, who demonstrated that only under UV light illumination the decomposition of organic contaminants can take place leading to the creation of superhydrophilic surfaces.<sup>11</sup> In another studies, Munuera *et al.* and Anpo *et al.* have reported that H<sub>2</sub>O desorbs from the TiO<sub>2</sub> surface during UV light irradiation, i.e., photodesorption of H<sub>2</sub>O, due to the heating effect of the light source.<sup>14,15</sup> Recently, special attention has been focused on the particle network before and after irradiation and how it changes during illumination.<sup>16,17</sup> Wang *et al.* reported that the metal oxide nanoparticles tend to form 3D networks via a self-aggregation mechanism when suspended in aqueous solution.<sup>16</sup> Some investigations indicated that UV light activation of TiO<sub>2</sub> leads to an increase in the surface area due to de-aggregation of particles agglomerates which in turn enhances the photonic efficiency.<sup>18,17</sup> The behaviours of the water adsorption and the associated hydroxyl groups during irradiation are, however still not well understood. The relative roles of the active oxygen and the hydroxyl groups for the photocatalytic process and for the photoinduced hydrophilicity, which have been under debate for a long time, still need to be investigated in detail. From this point of view, one of the most important issues of the current investigations concerning photocatalytic reactions and photoinduced hydrophilicity is to

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† Electronic Supplementary Information (ESI) available. See DOI: 10.1039/x0xx00000x

understand the fundamental process of water adsorption and the source of hydroxyl groups formed on TiO<sub>2</sub> surface under UV irradiation.

Herein, a detailed investigation of the H<sub>2</sub>O and/or D<sub>2</sub>O adsorption on the TiO<sub>2</sub> surface is reported using Attenuated Total Reflection Fourier Transformed Infrared spectroscopy (ATR-FTIR). Qualitative and quantitative evaluations of the D<sub>2</sub>O effect and the isotopic exchange on the TiO<sub>2</sub> surface are presented. This work focuses on the analysis of the behaviours of OH and OD stretching groups before and after UV irradiation considering the role of molecular oxygen to provide new insights into the mechanism of photoinduced hydrophilicity during UV illumination.

## 2. Experimental

### 2.1 Materials

The photocatalyst, titanium dioxide PC-500 (Cristal Global) consisted of 100% pure anatase with a specific BET-surface area of 340 m<sup>2</sup> g<sup>-1</sup> and a primary particle size of 5-10 nm. Deuterium oxide (D<sub>2</sub>O) (99.9 atom % D) was purchased from Sigma Aldrich. Deionized water (H<sub>2</sub>O) was supplied from a Millipore Mill-Q system with a resistivity equal to 18.2 Ω cm at 25 °C.

### 2.2 ATR-FTIR Spectroscopic Measurements

#### 2.2.1 Preparation of TiO<sub>2</sub> Films on the ATR-FTIR Crystal

An aqueous suspension of TiO<sub>2</sub> at a concentration of 5.75 g l<sup>-1</sup> was initially prepared and sonicated for 15 min in an ultrasonic cleaning bath. An aliquot of 400 μL of the TiO<sub>2</sub> suspension was placed on the surface of the ZnSe ATR crystal. This small volume was simply spread by balancing the unit manually. The suspension was then directly allowed to evaporate to dryness by storing the crystal in a semi-opened desiccator at room temperature. Prior to deposition of the TiO<sub>2</sub> films, the ZnSe surface (area = 6.8 mm×72 mm) was cleaned by polishing with 1 mm diamond paste (Metadi II, polishing grade) and rinsed with methanol and deionised water. The coverage of the final dry layer of particles obtained was 2.3 g m<sup>-2</sup> and the layer appeared homogeneous under visual inspection. In the original preparation by Hug *et al.*, AFM measurements of layers with coverage of 2.3 g m<sup>-2</sup> yielded a thickness of 1-3 μm.<sup>19,20</sup> The final resulting layers of particles remained stable over the entire

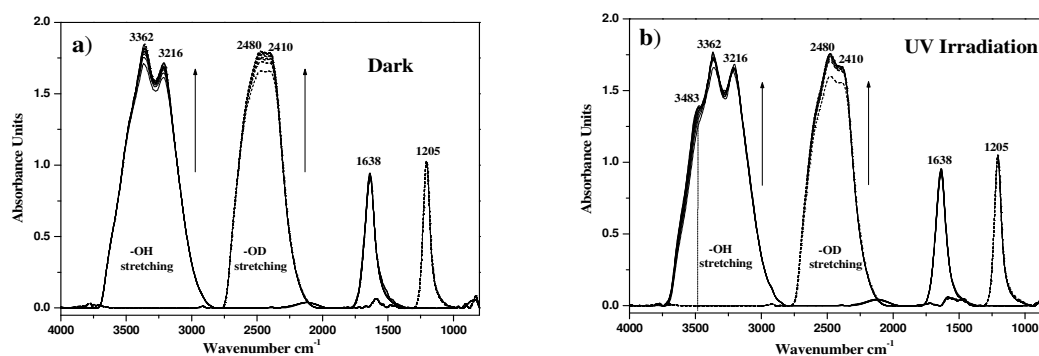
course of the experiment. Thus, it can be assumed that the effective path lengths at all wavelengths remained unchanged.

#### 2.2.2 ATR-FTIR Measurement

The ATR-FTIR spectra of the TiO<sub>2</sub> samples were monitored by a FTIR spectrometer (IFS 66 BRUKER) equipped with an internal reflection element 45° ZnSe crystal and a deuterated triglycine sulfate (DTGS) detector. The interferometer and the infrared light path in the spectrometer were constantly purged with Argon to avoid H<sub>2</sub>O and CO<sub>2</sub> contamination. The spectra were recorded with 350 scans at 4 cm<sup>-1</sup> resolution and analysed using OPUS version 6.5 software. The background was subtracted and baseline correction was made in order to eliminate minor fluctuations due to instrumental instability. Irradiation of samples with UV(A) light was carried out using an LED lamp (Model LED-Driver, THORLABS) emitting UV light (365 nm). The distance from the UV lamp to the surface of the test solution was kept at 30 cm on which the intensity of UV(A) light was of 1.0 mWcm<sup>-2</sup> as measured by an UV radiometer (Dr. Hönl GmbH, Martinsried, Germany). Prior to starting the irradiation experiments, spectra of adsorption of H<sub>2</sub>O and D<sub>2</sub>O on the TiO<sub>2</sub> were monitored in the dark. When the last spectrum of each experiment had been recorded, the UV(A) lamp was turned on and another sequence of spectra was recorded. These two groups of spectra were considered as the blank reference spectra in the dark and under UV(A) illumination, respectively.

## 3. Results

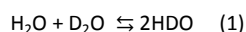
Fig. 1a shows the time evolution of the spectra of adsorbed H<sub>2</sub>O and D<sub>2</sub>O on TiO<sub>2</sub> in the dark. The signal of the adsorbed water on TiO<sub>2</sub> with maxima at 3362, 3216, and 1638 cm<sup>-1</sup> can be clearly seen. As is well known, the adsorption of H<sub>2</sub>O is represented by strong IR absorbances of O-H stretching in the region between 3200 and 3550 cm<sup>-1</sup> and δ (H-O-H) bending at 1638 cm<sup>-1</sup> which is assigned to undissociated water molecules. When D<sub>2</sub>O was used instead of H<sub>2</sub>O, the results showed that all bands of adsorbed water on TiO<sub>2</sub> surface were shifted to lower frequency with exchanging H for D. Additionally, an increase in the intensities of OH and OD stretching has been observed by increasing the adsorption time in the dark, which can be explained by increased adsorption of water molecules as well as of deuterium oxide on the TiO<sub>2</sub> surface until attaining an equilibrium.



**Fig. 1** Time evolution of the ATR-FTIR spectra of adsorbed pure H<sub>2</sub>O (solid lines) and D<sub>2</sub>O (dashed lines) in the presence of O<sub>2</sub> on TiO<sub>2</sub> a) in the dark for 4 h, b) under 8 h of UV(A) illumination.

It can be clearly seen from Fig. 1b that subsequent illumination with UV(A) light in the presence of O<sub>2</sub>, results in the formation of a shoulder in the FTIR spectrum of the adsorbed H<sub>2</sub>O at 3483 cm<sup>-1</sup> as well as an increase in the bands at 2480 cm<sup>-1</sup> in the case of D<sub>2</sub>O (Fig. 1b). The shoulder formed around 3483 cm<sup>-1</sup> can be attributed to the photocatalytically generated hydroperoxide. This was confirmed by recording the spectrum of H<sub>2</sub>O<sub>2</sub> added to the water-TiO<sub>2</sub> system in the dark in which a similar band (at 3483 cm<sup>-1</sup>) has been detected (Fig. S1, in supplementary Data). Furthermore, no formation of such band around 3483 cm<sup>-1</sup> was observed in the absence of O<sub>2</sub> (Fig. S2, in supplementary Data).

Fig. 2 shows the spectra of D<sub>2</sub>O–H<sub>2</sub>O mixtures adsorbed on TiO<sub>2</sub> in the dark. It is obvious from these spectra that the intensity of the band in the OH-stretching region decreased gradually with increasing dosage of D<sub>2</sub>O. Interestingly, the peak of the isotopologue HDO bending band centred at 1450 cm<sup>-1</sup> was formed and increased by increasing the concentration of H<sub>2</sub>O in D<sub>2</sub>O until it approached equimolar proportions (Fig.2 insert). This peak then decreased again with further increases in the amounts of H<sub>2</sub>O in D<sub>2</sub>O. The mixture of H<sub>2</sub>O and D<sub>2</sub>O reacted and equilibrated via transfer by self-dissociation and recombination to form HDO.<sup>21</sup>



The equilibrium constant K of this reaction is given by

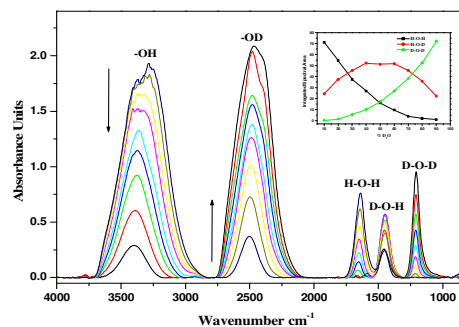


Fig. 2 ATR-FTIR spectra of D<sub>2</sub>O–H<sub>2</sub>O mixtures with different concentrations adsorbed on TiO<sub>2</sub> in the dark.

$$K = [\text{HOD}]^2/[\text{H}_2\text{O}][\text{D}_2\text{O}] \quad K_{\text{eq}} = 3.86 \text{ at } 25^\circ\text{C} \quad (2)$$

To investigate the effect of D<sub>2</sub>O on the hydroxyl group behaviour during UV irradiation, a series of experiments for water adsorption with different ratios of D<sub>2</sub>O were performed under irradiation in the presence of oxygen. Fig. 3 shows the time evolution of the intensities of the integrated spectral areas of OH and OD stretching groups that take place on the surface of TiO<sub>2</sub> before and after UV

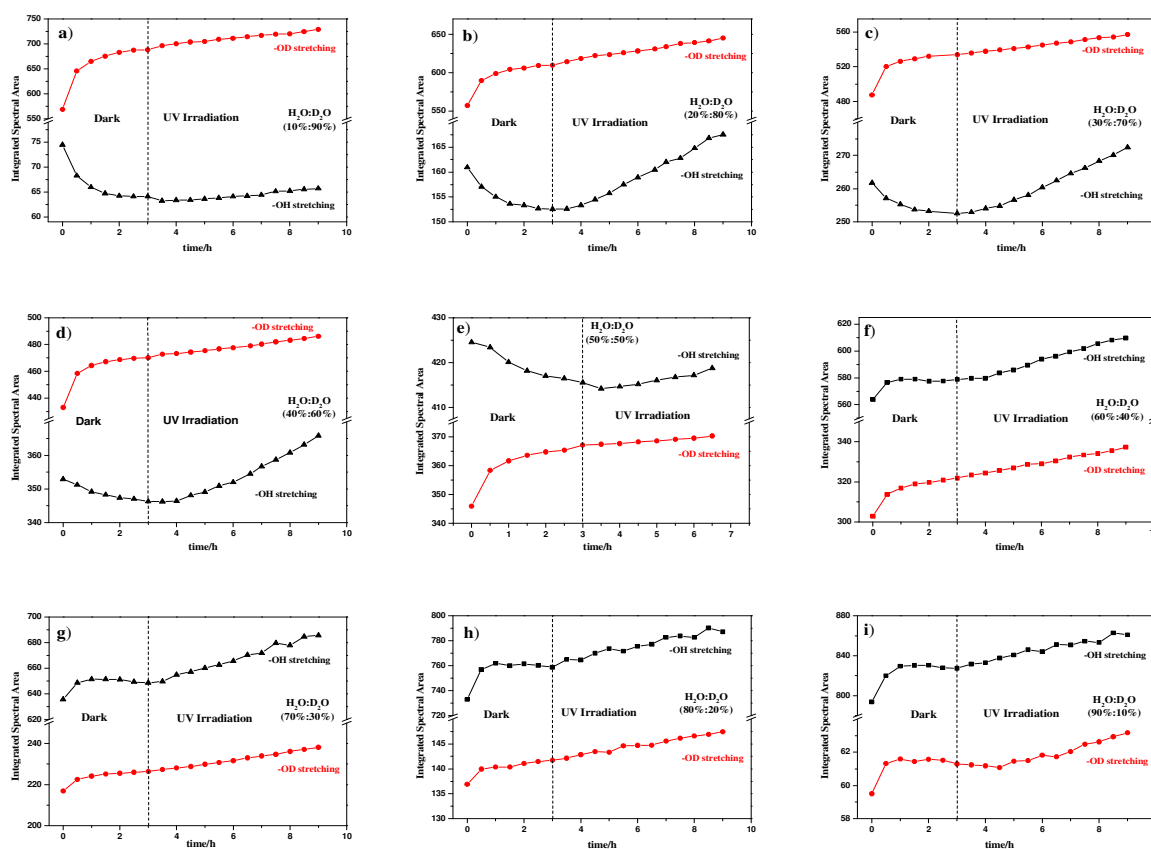


Fig. 3 Time evolution of the intensity of the integrated spectral areas of the OH and OD stretching groups before and after UV(A) irradiation with different ratios (H<sub>2</sub>O:D<sub>2</sub>O).

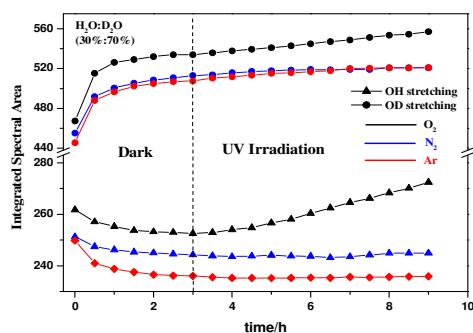
irradiation. As can be clearly seen, in the dark, at low concentrations of H<sub>2</sub>O (H<sub>2</sub>O% < D<sub>2</sub>O %), a strong decrease in the amount of OH stretching was observed especially in the first hour. Simultaneously, the amount of OD stretching groups adsorbed on TiO<sub>2</sub> increased during this period until the system reached equilibrium. In contrast, at higher concentrations of H<sub>2</sub>O (H<sub>2</sub>O% > D<sub>2</sub>O%), both OH and OD stretching bands increased. Interestingly, at an equimolar mixture of 50% H<sub>2</sub>O and D<sub>2</sub>O, a decrease in intensity of OH stretching with a simultaneous increase in the OD stretching could be observed in the dark. From these results we suggest that the deuteride ion in the dark shows a stronger adsorption ability than hydroxyl ions at the surface of TiO<sub>2</sub>. Thus, the deuteride ions could lead to isotopic exchange by replacing hydroxyl groups adsorbed on the TiO<sub>2</sub> surface (reaction 3):



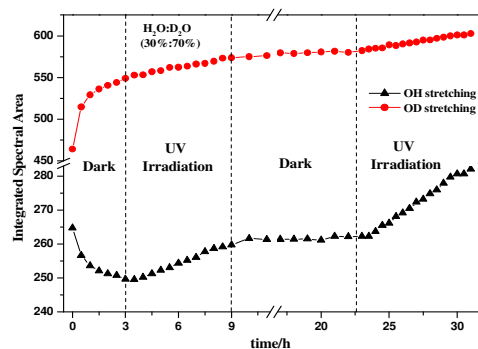
It is worth noting that the time required to reach equilibrium for all experiments was nearly constant, i.e., it took almost 3 hours to approach isotopic exchange equilibrium. Furthermore, this evaluation of OD groups could be associated with the better exchange of the isolated Ti-OH species on TiO<sub>2</sub> at higher concentrations than at lower concentrations of D<sub>2</sub>O. Interestingly, upon UV(A) irradiation, an increase in the intensity of the OH stretching could be observed almost immediately as well as an increasing intensity of the OD stretching band, i.e., Fig. S3 (see supplementary Data). These results, however can be largely attributed to either an increase of the amount of OH and OD groups by the formation of new OH and OD groups on the surface or by increasing the amount of H<sub>2</sub>O and D<sub>2</sub>O molecules chemisorbed on the TiO<sub>2</sub> surface.

To determine the effect of molecular oxygen on the behaviour of the H<sub>2</sub>O and D<sub>2</sub>O adsorption on the TiO<sub>2</sub> surface, the solution was saturated and constantly purged with different types of gases, i.e., oxygen, nitrogen, and argon during the ATR-FTIR measurements. As shown in Fig. 4, the integrated intensity of the OH and OD stretching group increased significantly upon illumination in the presence of molecular O<sub>2</sub>. By contrast, when the sample was purged with nitrogen or argon, no increase in the OH and/or OD stretching group was observed. A similar isotopic exchange behaviour, however, was observed during the dark period.

The time evolution of the intensity of the integrated spectral areas



**Fig. 4** Evolution of the intensity of the integrated spectral areas of OH and OD stretching groups before and after UV irradiation: Effect of dissolved O<sub>2</sub>, N<sub>2</sub> and Ar on the adsorption of H<sub>2</sub>O-D<sub>2</sub>O on the TiO<sub>2</sub> surface.



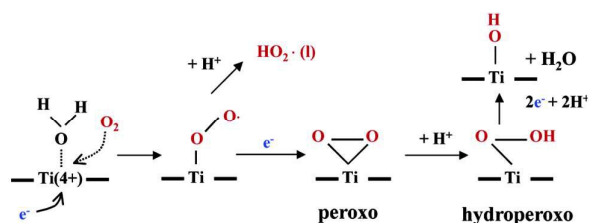
**Fig. 5** Time evolution of the intensity of the integrated spectral areas of OH and OD stretching groups before and after UV(A) irradiation with prolonged time (Alternately).

of the OH and OD stretching groups before and after UV irradiation at prolonged time periods are shown in Fig. 5. Again as described above under dark conditions, the isotopic exchanges inevitably took place on the TiO<sub>2</sub> surface. When the system was subsequently stored again in the dark (after the UV illumination was turned off), however, no more isotopic exchange behaviour was detected and the amount of OH and OD groups remained steady during this period. Additionally, when the system was illuminated again, the amount of OH and OD groups immediately increased.

#### 4. Discussion

When the TiO<sub>2</sub> surface was illuminated with UV light, electron and hole pairs were generated, and they, respectively, reduced and oxidized adsorbates on the surface. As can be seen from the ATR-FTIR spectra (Fig. 1b), a clear shoulder appears at 3483 cm<sup>-1</sup> upon irradiation of the system by UV(A) in the presence of molecular oxygen (O<sub>2</sub>). This peak can most likely be assigned to a Ti-OOH vibration as has been confirmed from the spectrum of H<sub>2</sub>O<sub>2</sub> added to the water-TiO<sub>2</sub> system in the dark in which a similar band (at 3483 cm<sup>-1</sup>) has been detected as shown in Fig. S1 (see supplementary Data). The formation of this adsorbed hydroperoxy species only in the presence of (O<sub>2</sub>) can be considered as evidence that it is resulting from the reduction of molecular oxygen adsorbed at the TiO<sub>2</sub> surface by the photogenerated electrons rather than from the oxidation of water by the holes. These results are in agreement with Nakamura *et al.*, who revealed and assigned O-O stretching of a surface peroxo species to identify Ti-OOH as a primary intermediate from the O<sub>2</sub> photoreduction.<sup>22</sup> Robertson *et al.*, however, reported that the formation of hydroxyl radicals for the photocatalytic process could be preferable through the water oxidation at the valence band rather than through the oxygen reduction at the conduction band.<sup>23,6</sup> Based on the identification of these intermediates, a mechanism for the photocatalytic formation of hydroperoxide is proposed (Scheme 1).<sup>22,24</sup> In this mechanism, the photogenerated conduction band electrons reduce Ti<sup>4+</sup> at the surface that adsorbs H<sub>2</sub>O, then molecular oxygen attacks immediately to form superperoxo TiOO<sup>•</sup>. This superperoxo is reduced to peroxo Ti(O<sub>2</sub>), which is equivalent to the hydroperoxide Ti-OOH when it is protonated. Furthermore, an increase in the intensities of the OH and OD stretching bands is observed after illumination which may be assigned either to the stretching modes

**Scheme 1** Proposed mechanism for the photocatalytic formation of hydroperoxides. (Copyright 2003 American Chemical Society.)



of Ti-OH formed upon illumination or to the OH stretching vibration of adsorbed water molecules which in turn increases the hydrophilicity of the TiO<sub>2</sub> surface. Hashimoto *et al.* reported that during UV irradiation not only the photocatalytic reactions take place on the surface but also the photoinduced hydrophilicity resulting from the separation and diffusion of photogenerated electrons and holes occurring on the surface of TiO<sub>2</sub>.<sup>3</sup> As shown in Fig. 3, at low concentrations of H<sub>2</sub>O (H<sub>2</sub>O% < D<sub>2</sub>O %), the reduction of the OH band with the simultaneous increase in the OD stretching band in the dark revealed that the deuteride ions have a stronger adsorption affinity than the hydroxyl ions. Furthermore, it can be clearly observed that the photoinduced hydrophilicity of the TiO<sub>2</sub> surfaces which was explained by an increase in the amount of hydroxyl groups is readily induced upon UV light irradiation in the presence of O<sub>2</sub>. These results can be largely attributed to either an increase in the number of surface OH groups or to an increased adsorption of water molecules on the TiO<sub>2</sub> surface.

Several mechanisms suggested that the photoinduced holes play a crucial role in the photoinduced hydrophilicity conversion. It was thus proposed that the photogenerated holes produced in the bulk of TiO<sub>2</sub> diffuse to the surface where they are trapped at lattice oxygen sites. The trapped holes are consumed by reaction with the TiO<sub>2</sub> itself creating oxygen vacancies,<sup>25</sup> or by breaking the bond between the lattice titanium and oxygen ions upon the coordination of water molecules at the titanium site.<sup>10</sup> Subsequently the water molecule dissociates in an oxygen vacancy or releases proton for charge compensation to produce new OH groups. As shown in Fig. 5, however, an isotopic exchange between deuteride ions and hydroxyl groups takes place on the TiO<sub>2</sub> surface in the dark according to equation 3. After 6 h of irradiation even

**Scheme 2** Interaction model of adsorbed water during UV(A) light irradiation in the absence and presence of O<sub>2</sub> molecules.

during a prolonged time in the dark, however, no isotopic exchange was detected between hydroxyl groups and deuteride ions. From this result we suggest that the increase of OH and OD stretching groups under irradiation is most likely caused by adsorption of H<sub>2</sub>O and D<sub>2</sub>O molecules, respectively, and not by the generation of hydroxyl groups. There have been several reports supporting our suggestion that water molecules strongly adsorb onto the TiO<sub>2</sub> surface under UV illumination. Nosaka *et al.* reported that the increase in the amount of adsorbed water on the TiO<sub>2</sub> surface after UV light irradiation was observed by <sup>1</sup>H NMR spectroscopy.<sup>26</sup> Another study by Uosaki *et al.* using SFG spectroscopy confirmed that the UV illumination increased not only the amount of adsorbed water but also the order of the adsorbed water on the TiO<sub>2</sub> surface.<sup>27</sup> Although the mechanism for the H<sub>2</sub>O/UV light interactions still remains uncertain, it seems that the particle network during UV light irradiation might be responsible for the enhancement of water adsorption on a TiO<sub>2</sub> particle film due to a new distribution of the particles. Wang *et al.* revealed the fact that under UV(A) illumination the total exposed TiO<sub>2</sub> surface increases due to the de-aggregation of particles agglomerates which was explained by assuming that part of the absorbed light energy is converted non-adiabatically into heat which is subsequently used to break the bonds between the particles thus producing additional surface area for the photocatalytic process.<sup>17,28</sup> One of the first ATR-FTIR studies that focused on the adsorption phenomena in aqueous solution was performed by Mendive *et al.*, who combined experimental data and theoretical calculations to demonstrate that the water molecule can fill the space in between the particles which was demonstrated by the increase in the IR band corresponding to the bending mode of water.<sup>29</sup>

We propose that UV irradiation leads to an increase in the amount of adsorbed water by thermal processes with hydrophilicity effects simultaneously taking place on the TiO<sub>2</sub> surface. Moreover, as shown in Fig. 4, the results clearly indicate that the presence of O<sub>2</sub> is necessary to enhance the photoadsorption of H<sub>2</sub>O and D<sub>2</sub>O on TiO<sub>2</sub> surfaces during UV(A) irradiation. This confirms that the adsorption of water and D<sub>2</sub>O to TiO<sub>2</sub> surfaces is caused by a photoinduced charge transfer process in presence of molecular oxygen. Takeuchi *et al.*<sup>30</sup> proposed a mechanism which revealed that when TiO<sub>2</sub> surfaces were irradiated with UV light in the absence of O<sub>2</sub>, the electrons trapped on the Ti<sup>3+</sup> sites were not scavenged by O<sub>2</sub> and the holes trapped on the TiO<sub>2</sub> surface were immediately consumed to oxidize the lattice oxygen, resulting in the formation of oxygen vacancies. Such photoreduced TiO<sub>2</sub> surfaces can behave as negatively charged surfaces. Thus, water molecules could hardly adsorb on the photoreduced surfaces due to repulsive effects (Scheme 2). On the other hand, when the TiO<sub>2</sub> surfaces are irradiated in the presence of O<sub>2</sub>, the surface trapped electrons are scavenged by O<sub>2</sub> and the photoreduced surfaces are immediately oxidized. In this way, the TiO<sub>2</sub> surface is very likely interacting with water molecules due to interaction effects resulting in an increase in the amount of adsorbed water. The results of this investigation indicate that the photoinduced hydrophilic effect is achieved by an increase in the amount of adsorbed H<sub>2</sub>O molecules and this phenomenon occurs not only by a thermal process but also by a photoinduced charge transfer process.

## 5. Conclusions

The behaviour of adsorbed H<sub>2</sub>O and D<sub>2</sub>O in photocatalytic processes and in photoinduced hydrophilicity on TiO<sub>2</sub> surfaces has been addressed using ATR-FTIR spectroscopy. This study made it possible,

to investigate adsorption processes at the TiO<sub>2</sub>/aqueous solution interface, quantitatively and in-situ. It was shown that the photocatalytic reduction of O<sub>2</sub> at the anatase surface leads to the formation of hydroperoxo groups as demonstrated by a FTIR band at 3483 cm<sup>-1</sup>. Adsorption of H<sub>2</sub>O with different ratios of D<sub>2</sub>O on TiO<sub>2</sub> revealed that the deuteride ions exhibit stronger adsorption ability than hydroxyl ions resulting in an isotopic exchange reaction which takes place in the dark. Upon illumination with UV light in the presence of O<sub>2</sub> both OH and OD groups are formed, which in turn increase the hydrophilicity of the TiO<sub>2</sub> surface. In contrast, when the solution was saturated with argon or nitrogen, no formation of OH and OD groups was detected and the hydrophilicity was strongly inhibited. The increase in the amount of OH and OD groups is suggested to be caused by photoadsorption of H<sub>2</sub>O and D<sub>2</sub>O due to de-aggregation of the pre-existing network of particles rather than by dissociative water and D<sub>2</sub>O adsorption at surficial oxygen vacancies. Finally, it is concluded that the role of oxygen during UV irradiation is very crucial for the presence of both the photoinduced hydrophilic as well as the photocatalytic response.

### Acknowledgements

Belhadj H. gratefully acknowledges a scholarship from the Deutscher Akademischer Austauschdienst (DAAD) providing the financial support to perform his Ph.D. studies in Germany. The present study was performed within the Project "Establishment of the Laboratory 'Photoactive Nanocomposite Materials'" No. 14.Z50.31.0016 supported by a Mega-grant of the Government of the Russian Federation.

### References

- M. R. Hoffmann, M. R. Hoffmann, S. T. Martin, S. T. Martin, W. Choi, W. Choi, D. W. Bahnemann and D. W. Bahnemann, *Chemical Reviews*, 1995, **95**, 69–96.
- J. Tschirch, R. Dillert, D. Bahnemann, B. Proft, A. Biedermann and B. Goer, *Research on Chemical Intermediates*, 2008, **34**, 381–392.
- M. Miyauchi, A. Nakajima, T. Watanabe and K. Hashimoto, *Chemistry of Materials*, 2002, **14**, 2812–2816.
- J. Xu, N. Sahai, C. M. Eggleston and M. a. a. Schoonen, *Earth and Planetary Science Letters*, 2013, **363**, 156–167.
- T. Hirakawa and Y. Nosaka, *Langmuir*, 2002, **18**, 3247–3254.
- P. K. J. Robertson, D. W. Bahnemann, L. a. Lawton and E. Bellu, *Applied Catalysis B: Environmental*, 2011, **108-109**, 1–5.
- R. Wang, K. Hashimoto, A. Fujishima, M. Chikuni, E. Kojima, A. Kitamura, M. Shimohigoshi and T. Watanabe, *Advanced Materials*, 1998, **10**, 135–138.
- L. Zhang, R. Dillert, D. Bahnemann and M. Vormoor, *Energy & Environmental Science*, 2012, **5**, 7491.
- L. M. Liu, P. Crawford and P. Hu, *Progress in Surface Science*, 2009, **84**, 155–176.
- N. Sakai, A. Fujishima, T. Watanabe and K. Hashimoto, *The Journal of Physical Chemistry B*, 2003, **107**, 1028–1035.
- T. Zubkoy, D. Stahl, T. L. Thompson, D. Panayotov, O. Diwald and J. T. Yates, *Journal of Physical Chemistry B*, 2005, **109**, 15454–15462.
- Hiroshi Irie, Kazuhito Hashimoto, *Environmental Photochemistry Part II*, 2005, vol. 2M.
- K. Hashimoto, H. Irie and A. Fujishima, *Japanese Journal of Applied Physics*, 2005, **44**, 8269–8285.
- G. Munuera, A. R. Gonzalez-Elipe, J. Soria and J. Sanz, *Journal of the Chemical Society, Faraday Transactions 1*, 1980, **76**, 1535.
- M. Takeuchi, K. Sakamoto, G. Martra, S. Coluccia and M. Anpo, *Journal of Physical Chemistry B*, 2005, **109**, 15422–15428.
- C. Wang, C. Böttcher, D. W. Bahnemann and J. K. Dohrmann, *Journal of Materials Chemistry*, 2003, **13**, 2322.
- C. Wang, R. Pagel, J. K. Dohrmann and D. W. Bahnemann, *Comptes Rendus Chimie*, 2006, **9**, 761–773.
- C. Y. Wang, R. Pagel, D. W. Bahnemann and J. K. Dohrmann, *Journal of Physical Chemistry B*, 2004, **108**, 14082–14092.
- S. J. Hug and B. Sulzberger, *Langmuir*, 1994, **10**, 3587–3597.
- C. B. Mendive, T. Bredow, A. Feldhoff, M. A. Blesa and D. Bahnemann, *Physical chemistry chemical physics: PCCP*, 2009, **11**, 1794–808.
- J. C. Duplan, L. Mahi and J. L. Brunet, *Chemical Physics Letters*, 2005, **413**, 400–403.
- R. Nakamura, A. Imanishi, K. Murakoshi and Y. Nakato, *Journal of the American Chemical Society*, 2003, **125**, 7443–50.
- L. A. Lawton, P. K. J. Robertson, B. J. P. A. Cornish, I. L. Marr and M. Jaspars, *Journal of Catalysis*, 2003, **213**, 109–113.
- G. Mattioli, F. Filippone and A. A. Bonapasta, *Journal of the American Chemical Society*, 2006, **128**, 13772–13780.
- R. Wang, K. Hashimoto, A. Fujishima, M. Chikuni, E. Kojima, A. Kitamura, M. Shimohigoshi and T. Watanabe, *Nature*, 1997, **388**, 431–432.
- A. Y. Nosaka, E. Kojima, T. Fujiwara, H. Yagi, H. Akutsu and Y. Nosaka, *The Journal of Physical Chemistry B*, 2003, **107**, 12042–12044.
- K. Uosaki, T. Yano and S. Nihonyanagi, *The Journal of Physical Chemistry B*, 2004, **108**, 19086–19088.
- J. Schneider, M. Matsuoka, M. Takeuchi, J. Zhang, Y. Horiuchi, M. Anpo and D. W. Bahnemann, *Chemical Reviews*, 2014, **114**, 9919–9986.
- C. B. Mendive, D. Hansmann, T. Bredow and D. Bahnemann, *Journal of Physical Chemistry C*, 2011, **115**, 19676–19685.
- M. Takeuchi, G. Martra, S. Coluccia and M. Anpo, *Journal of Physical Chemistry C*, 2007, **111**, 9811–9817.