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## COMMUNICATION

# Room-Temperature Acetylene Hydration by a Hg(II)-laced Metal-Organic Framework<sup>†</sup>

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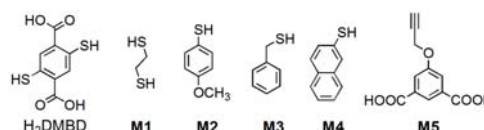
**Thiol (-SH) groups within a Zr(IV)-based metal-organic framework (MOF) anchor Hg(II) atoms, which was then oxidized to acidic sulfonate functions for catalyzing acetylene hydration at room temperature.**

Among the various ways to functionalize the porous solids of metal-organic frameworks (MOFs),<sup>1</sup> the versatile reactivity of the thiol (-SH) group offers unique advantages.<sup>2</sup> For example, thiols as strong soft donors readily take up various metal ions, which closely bears on the removal of heavy metal ions,<sup>3</sup> and on the creation of electroactive/semiconducting<sup>2b, 4</sup> or catalytic sites<sup>5</sup> (e.g., mimicking the iron-sulfur, copper-sulfur proteins) within the MOF matrices. Notably, recent exercises (e.g., using 2,5-dimercapto-1,4-benzenedicarboxylic acid, H<sub>2</sub>DMBD, chart 1) indicated that dense arrays of free-standing thiol groups can be built into the host net when chemically very hard ions [like Eu(III), Zr(IV) or Al(IV); i.e., these tend to stay unbonded with the thiol group] are chosen to link up the carboxyl groups.<sup>3a, 5c</sup>

As part of our ongoing efforts to further exploit the thiol/thiolate groups thus installed within MOF solids, we here utilize simple oxidation to effect the conversion into sulfonic acid and metal sulfonate functions.<sup>6</sup> Such conversion is intended to liberate the proton and metal centers from the thiol groups, and to create strong acidity and reactivity properties within the MOF pores.

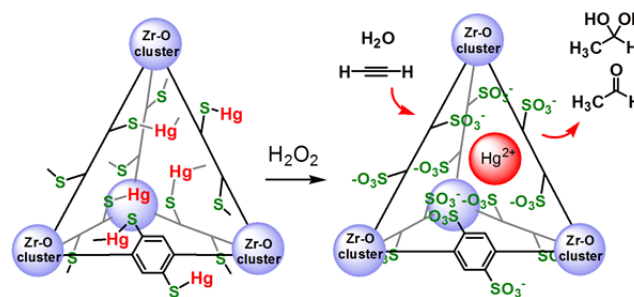
One major advantage of this method lies in the dense array of sulfonate units that can be installed (e.g., two per linker, as from

DMBD). Previously, MOF systems (e.g., MIL-101(Cr), MIL-53(Al)<sup>7</sup> and others<sup>8</sup>) had been directly sulfated (e.g., by ClSO<sub>3</sub>H);



**Chart. 1** Molecule H<sub>2</sub>DMBD for building the ZrDMBD solid, **M1-M4** for the selective uptake test with the ZrBDSO<sub>3</sub>-Hg solid, and **M5** for a hydration test benchmarked against acetylene.

but the sulfonate group, once attached, deactivated the aromatic core and thus hindered further sulfation (i.e., the number of installed sulfonate is limited). In another approach, sulfated ligands and pristine ligands as a mixture were reacted directly with metal ions to form the framework, but potential interference from the sulfonate group in binding with the metal ions (and thus disrupting the network construction) often limit the fraction of the sulfonated ligands. More broadly, the oxidation of the metal thiolate moiety generates *in situ* metal sulfonate functions on the host net, whereas for other sulfated frameworks, additional steps of ion exchange are necessary for inserting exo-framework metal ions.<sup>9</sup> As thiol groups readily bind various metal ions, our approach offers flexible control over the amount and type of metal ions to be deployed in the pores.



**Fig. 1** Schematics for H<sub>2</sub>O<sub>2</sub> oxidation of the ZrDMBD-Hg net (left) into the ZrBDSO<sub>3</sub>-Hg net (right) and the latter's use in catalysing the hydration of acetylene. The host net is simplified as a tetrahedral cage, with each Zr<sub>6</sub>O<sub>4</sub>(OH)<sub>4</sub> cluster shown as purple spheres.

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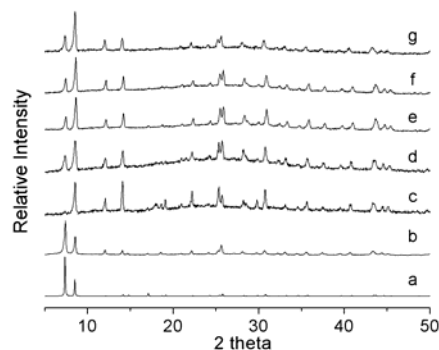
<sup>†</sup> Dedicated to the 60th birthday of Prof. Stephen Lee.

Electronic Supplementary Information (ESI) available: [Experimental procedures; network synthesis and activation; SEM photographs; elemental analysis, CO<sub>2</sub> sorption, TGA, IR/Raman, and PXRD data]. See DOI: 10.1039/x0xx00000x

For illustration, we here present a MOF solid with  $\text{Hg}^{2+}$ -sulfonate functions as an especially active catalyst for the acetylene hydration reaction.

First we introduce the three major stages of sample preparation.

1) Reaction of  $\text{ZrCl}_4$  and  $\text{H}_2\text{DMBD}$  under solvothermal conditions yielded a crystalline powder sample of the thiol-laced framework  $\text{ZrDMBD}$  (similar procedure as reported,<sup>3a</sup> but with  $\text{N}_2$  protection to minimize the oxidation of the  $-\text{SH}$  groups). The composition of the  $\text{ZrDMBD}$  sample features a  $\text{Zr}_6\text{O}_4(\text{OH})_4(\text{DMBD})_6$  framework with DMF and  $\text{H}_2\text{O}$  guests (see SI for details; see also Fig. 1, left for a



**Fig. 2** X-ray powder patterns ( $\text{Cu K}\alpha = 1.5418 \text{ \AA}$ ) of (a) a simulation from a structure model of  $\text{ZrDMBD}$ ; (b) an as-made  $\text{ZrDMBD}$  sample; (c)  $\text{ZrDMBD-Hg}$ ; (d)  $\text{ZrBDSO}_3\text{-Hg}$ ; (e-g)  $\text{ZrBDSO}_3\text{-Hg}$  after the 1<sup>st</sup>, 2<sup>nd</sup> and 3<sup>rd</sup> cycle of acetylene hydration catalysis test, respectively.

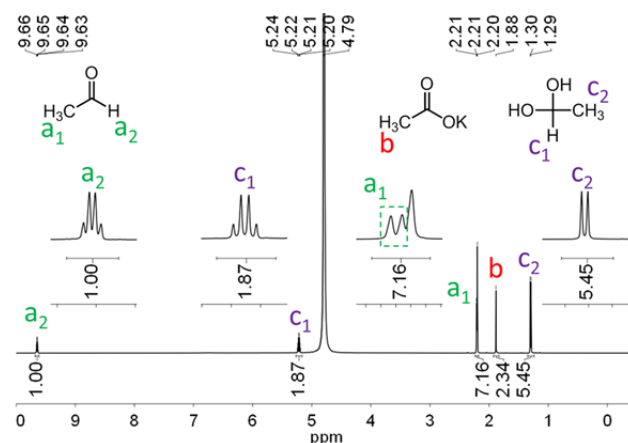
schematic of the framework). 2) Treatment of  $\text{ZrDMBD}$  with an aqueous solution of  $\text{HgCl}_2$  led to the mercurated solid  $\text{ZrDMBD-Hg}$ , which was found to contain a 1:6:4  $\text{Zr}_6\text{O}_4(\text{OH})_4/\text{DMBD}/\text{Hg}$  ratio (together with  $\text{Cl}^-$ , DMF and  $\text{H}_2\text{O}$ ; equivalent to w/w 21.4% for Hg; see SI for details). The substantial Hg presence in  $\text{ZrDMBD-Hg}$  was also revealed by the absence of S-H in the IR/Raman spectra (Fig. S1 and S2), as well as the large change in the intensity profile of the PXRD patterns (cf. patterns b and c in Fig. 2). 3) Oxidation by  $\text{H}_2\text{O}_2$  on  $\text{ZrDMBD-Hg}$  converts the thiolate groups into sulfonate functions; the resultant solid (denoted as  $\text{ZrBDSO}_3\text{-Hg}$ ) was found by elemental analyses (see SI) to feature the composition,  $\text{Zr}_6\text{O}_4(\text{OH})_4[\text{C}_8\text{H}_2\text{O}_4(\text{SO}_3^-)_{0.8}(\text{SO}_3\text{H})_{1.2}]_6 \cdot \text{Hg}_{2.4} \cdot (\text{H}_2\text{O})_{50}$ , with the formation of the sulfonate functions ( $-\text{SO}_3^-$ ) being indicated by the IR (Fig. S1) and NMR spectra (Fig. S3). In spite of the drastic oxidative transformation, the structural integrity of the host net was found by PXRD to be intact (Fig. 2, pattern d). Compared with the  $\text{ZrDMBD-Hg}$  sample, the Hg content in  $\text{ZrBDSO}_3\text{-Hg}$  (11.8%) is lower, i.e., about 40% of the Hg leached away during the  $\text{H}_2\text{O}_2$  treatment; however, such  $\text{Hg(II)}$  leaching from the solid host can be readily suppressed by using a  $\text{H}_2\text{O}_2$  solution containing dissolved  $\text{Hg}(\text{NO}_3)_2$ , e.g., with the resultant Hg content being 20.4% (see SI for the procedure and PXRD pattern g in Fig. S4). For the following catalytic study, the  $\text{ZrBDSO}_3\text{-Hg}$  sample (i.e., with a 6:2.4 linker/Hg ratio) was prepared by the simple  $\text{H}_2\text{O}_2$  treatment--without the added  $\text{Hg}(\text{NO}_3)_2$  solute.

Can one use ion exchange (i.e., with  $\text{ZrBDSO}_3\text{H}$ ) instead to access the  $\text{ZrBDSO}_3\text{-Hg}$  solid? To explore this possibility, the thiol groups in  $\text{ZrDMBD}$  were oxidized by  $\text{H}_2\text{O}_2$  into sulfonic acid groups--see SI for the procedure and characterization (Fig. S1-S4) on the resultant  $\text{ZrBDSO}_3\text{H}$  solid, and for the measured proton conductivity (Fig. S5; the conductivity can be improved with  $\text{H}_2\text{SO}_4$  treatment on the powder sample, as shown in a recent study<sup>6</sup>). Notably, ion exchange experiments on the  $\text{ZrBDSO}_3\text{H}$  solid thus obtained indicated lesser  $\text{Hg}^{2+}$  insertion. For example, even after the

$\text{ZrBDSO}_3\text{H}$  solid was heated in concentrated  $\text{Hg}(\text{NO}_3)_2$  and  $\text{HgCl}_2$  solutions for 18 hours (SI for the procedures and patterns d and e in Fig. S4), the Hg content in the solid was found by diphenylthiocarbazon extraction method analysis to be 2.8% and 3.0% (significantly lower than the values of 11.8%-20.4% in the  $\text{ZrBDSO}_3\text{-Hg}$  samples obtained from  $\text{H}_2\text{O}_2$  oxidation on  $\text{ZrDMBD-Hg}$ ). Such tests help to highlight  $\text{ZrDMBD-Hg}$  as an effective precursor to highly mercurated  $\text{ZrBDSO}_3\text{-Hg}$  products.

An additional test also helps to demonstrate that the  $\text{Hg}^{2+}$  ions are located inside of the  $\text{ZrBDSO}_3\text{-Hg}$  pores. Specifically, a mixture solution of four mercaptan molecules (in  $\text{CD}_2\text{Cl}_2$ ; **M1-M4** are shown in Chart 1) of increasing sizes was treated by  $\text{ZrBDSO}_3\text{-Hg}$  (containing Hg in excess relative to the thiols) at room temperature (rt). While NMR measurement indicated complete removal of the smallest **M1** (via the strong thiol-Hg interaction) from the solution within 12 hrs, the larger **M2**, **M3** and **M4** remained unchanged in concentrations (Fig. S6). Such size selectivity indicates that the  $\text{Hg}^{2+}$  ions are not accessible to the larger **M2-M4** mercaptans under these conditions, and points to potential applications in thiol uptake (see also Fig. S7 for the sorption test on 2-mercaptoethanol).

The catalytic efficacy of the  $\text{ZrBDSO}_3\text{-Hg}$  solid towards acetylene hydration of ( $\text{C}_2\text{H}_2$ ) was revealed in a simple reaction setup. Namely, by stirring at rt for a few hours a mixture of  $\text{ZrBDSO}_3\text{-Hg}$  solid (e.g., 100 mg, containing 0.059 mmol of Hg) and water (e.g., 2.7 mL; 0.15 mol) in a 1000-mL Schlenk flask filled with acetylene ( $\text{C}_2\text{H}_2$ ; atmospheric pressure; about 45 mmol, 1.2 g), an acetaldehyde content (in the form of acetaldehyde and the hydration product ethane-1,1-diol; Fig. 1) of 5.78 % (equivalent to a turnover number of 61 for Hg, Table 1) can be achieved in the water phase (supernatant; see Fig. 3 for the NMR spectrum). The product concentration compares well with the values (about 2-7%) normally produced in industrial reactor settings using homogeneous catalyst of  $\text{HgSO}_4/\text{H}_2\text{SO}_4$  solutions,<sup>10</sup> which nevertheless involve the more sophisticated conditions of constant flow of  $\text{C}_2\text{H}_2$  and significant heating (at 70-90°C). As a solid state catalyst,  $\text{ZrBDSO}_3\text{-Hg}$ , with



**Fig. 3**  $^1\text{H}$  NMR spectrum of the supernatant of acetylene reaction (dissolved in  $\text{D}_2\text{O}$  with  $\text{CH}_3\text{COOK}$  added as an internal standard). The internal standard peak at 1.88 ppm (singlet) and the products peak at 9.65 ppm (quartet) from acetaldehyde and 1.30 ppm (doublet) from ethane-1,1-diol were used to calculate the yield and TON.

both the  $\text{H}^+$  and  $\text{Hg}^{2+}$  agents lodged within its host het, provides the added advantages of non-corrosive (e.g., water) conditions and easy product isolation. Such advantages stand out even in comparison with the main-stream Wacker process,<sup>11</sup> wherein the highly

oxidizing and corrosive nature of the aqueous PdCl<sub>2</sub>/CuCl<sub>2</sub> catalyst (aggravated under the heated conditions) remains a concern.

For a more direct benchmark, a homogeneous catalyst consisting of HgSO<sub>4</sub> (18.2 mg, 0.061 mmol) dissolved in a H<sub>2</sub>SO<sub>4</sub> (18%, 2.7 mL) was examined under the same reaction conditions (e.g., 1.0 atm of C<sub>2</sub>H<sub>2</sub>, rt). Notice that both the Hg quantity and the solution volume are set to the same values as in the above test of the ZrBDSO<sub>3</sub>-Hg solid. In this homogeneous setting, the acetaldehyde products amounted to only 1.14% (TON: 12.8) in the solution, less than 1/5 of the value achieved by the ZrBDSO<sub>3</sub>-Hg solid catalyst. The efficiency of the ZrBDSO<sub>3</sub>-Hg solid catalyst can be ascribed to the compact arrangement of the H<sup>+</sup> and Hg<sup>2+</sup> agents within the host net, as well as to the hydrophobic aromatic struts promoting the C<sub>2</sub>H<sub>2</sub> diffusion into the pores.

**Table 1** Efficiencies of ZrBDSO<sub>3</sub>-Hg as a Lewis acid for hydration of acetylene.

$$\text{H}-\text{C}\equiv\text{C}-\text{H} + \text{H}_2\text{O} \xrightarrow[\text{rt}]{\text{catalyst}} \text{CH}_3\text{CHO} + \text{H}-\text{C}(\text{OH})_2-\text{H}$$

Catalyst	Hg content in MOF (wt%)	Total Hg (mg)	Product conc. (wt%)	TON <sup>[b]</sup>
ZrBDSO <sub>3</sub> -Hg cycle 1	11.8 <sup>[a]</sup>	11.8	5.78	61
ZrBDSO <sub>3</sub> -Hg cycle 2 <sup>[c]</sup>	11.9 <sup>[a]</sup>	8.3	5.95	62
ZrBDSO <sub>3</sub> -Hg cycle 3 <sup>[c]</sup>	11.4 <sup>[a]</sup>	6.9	5.68	59
HgSO <sub>4</sub> /H <sub>2</sub> SO <sub>4</sub>	N/A	12.3	1.14	13

<sup>[a]</sup> Mercury contents were determined by diphenylthiocarbazone extraction method. <sup>[b]</sup> TON is defined as the number of product formed per mercury atom. <sup>[c]</sup> See SI or the cycling procedure.

Also notably, the leaching of Hg(II) from the ZrBDSO<sub>3</sub>-Hg solid into the water phase (supernatant) is small: e.g., the supernatant (i.e., 2.7 mL) was found to contain only 0.11 mg mercury (equivalent to 40.7 ppm). In other words, less than 1% of the Hg content (11.8 mg) in the ZrBDSO<sub>3</sub>-Hg catalyst was leached into the water phase during the C<sub>2</sub>H<sub>2</sub> hydration process. The small Hg leaching, besides minimizing the environmental impact from the toxic Hg species, also makes it possible to recover the solid state catalyst for subsequent cycles of reactions--e.g., reducing the need for re-inserting Hg<sup>2+</sup> ions into the solid host.

The recovery involves oxidizing (back into Hg<sup>2+</sup>) the reduced Hg species resulted from side reactions. In the traditional homogeneous HgSO<sub>4</sub>/H<sub>2</sub>SO<sub>4</sub> systems, such side reactions were severe, forming large amount of Hg(0)/Hg(I)-containing sludge that had to be periodically drained from the industrial reactor. In the case of ZrBDSO<sub>3</sub>-Hg, as an indication of the ongoing reduction of Hg(II) ions, the white catalyst solid gradually developed a grey color, with concomitant decrease in catalytic activity. The used ZrBDSO<sub>3</sub>-Hg solid can be reactivated simply by immersion in a mixed solution of H<sub>2</sub>O<sub>2</sub>, HNO<sub>3</sub>, Hg(NO<sub>3</sub>)<sub>2</sub>·H<sub>2</sub>O at rt (e.g., for 15 minutes; see SI for details). The color of the solid returned to white, and the Hg content in the regenerated ZrBDSO<sub>3</sub>-Hg solid was found to be 11.9% (cf. 10.0% in the used catalyst), indicating the efficacy of the solid host in retaining the Hg(II) guests. The ZrBDSO<sub>3</sub>-Hg catalyst thus regenerated retains the structural integrity of the host lattice (e.g., see PXRD patterns e-g in Fig. 2) and continues to be highly active for

C<sub>2</sub>H<sub>2</sub> hydration, with turnover numbers (e.g., about 62) comparable to first round (see Table 1).

To demonstrate that the catalysis takes places within the pores of the ZrBDSO<sub>3</sub>-Hg solid, we examine the reactivity of a larger substrate, 5-propargyloxyisophthalic acid (**M5**), which, with a cross-section above 7 Å, is too bulky to enter into the host net (pore opening ~5 Å). To promote the solubility, a 2:1 THF/H<sub>2</sub>O solvent was used for the hydration reaction. No reaction was observed after stirring at rt for up to 8 hours a mixture of the ZrBDSO<sub>3</sub>-Hg solid and the THF/H<sub>2</sub>O solution of **M5** (see SI for details including NMR and TLC monitoring, e.g., Fig. S8 and S9); by contrast, when the homogeneous system of HgSO<sub>4</sub>/H<sub>2</sub>SO<sub>4</sub> was used instead, the homogeneous condition led to complete hydration of **M5** (e.g., also at rt and within 8 hrs; see NMR spectrum C of Fig. S8). This observation suggests that catalytic activity of ZrBDSO<sub>3</sub>-Hg entails substrates penetrating the host net, and the Hg<sup>2+</sup> ions operate from within the pores of the host net.

To sum up, the thiol function in ZrDMBD proves especially useful for accessing the Hg<sup>2+</sup>-laden solid of ZrBDSO<sub>3</sub>-Hg. The mild conditions (at rt, in water) for C<sub>2</sub>H<sub>2</sub> hydration attests to the enhanced activity of ZrBDSO<sub>3</sub>-Hg as a solid state catalyst. The catalytic activity likely results from the conjoint workings of the Hg<sup>2+</sup> and -SO<sub>3</sub>H acid units that are densely arrayed within the pores of the host net--e.g., neither ZrBMBD-Hg nor ZrBDSO<sub>3</sub>H exhibited observable catalytic activity under similar conditions (Fig. S14). We are working to access similar MOF materials with larger pores, in order to widen the scope of alkyne hydration applications.<sup>12</sup> Imbedding metal ions in porous frameworks offer great potential for opening novel reactivities, and thiol-laced frameworks will remain uniquely important in these studies.

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## Notes and references

- a) K. Manna, T. Zhang, F. X. Greene and W. Lin, *J. Am. Chem. Soc.*, 2015, **137**, 2665; b) S. Yuan, W. Lu, Y.-P. Chen, Q. Zhang, T.-F. Liu, D. Feng, X. Wang, J. Qin and H.-C. Zhou, *J. Am. Chem. Soc.*, 2015, **137**, 3177; c) J. Jiang, F. Gandara, Y.-B. Zhang, K. Na, O. M. Yaghi and W. G. Klemperer, *J. Am. Chem. Soc.*, 2014, **136**, 12844; d) J. M. Falkowski, T. Sawano, T. Zhang, G. Tsun, Y. Chen, J. V. Lockard and W. Lin, *J. Am. Chem. Soc.*, 2014, **136**, 5213; e) H. J. Jeon, R. Matsuda, P. Kanoo, H. Kajiro, L. Li, H. Sato, Y. Zheng and S. Kitagawa, *Chem. Commun.*, 2014, **50**, 10861; f) M. Inukai, S. Horike, W. Chen, D. Uneyama, T. Itakura and S. Kitagawa, *J. Mater. Chem. A*, 2014, **2**, 10404; g) J. Cui, Y.-L. Wong, M. Zeller, A. D. Hunter and Z. Xu, *Angew. Chem., Int. Ed.*, 2014, **53**, 14438; h) D. T. Genna, A. G. Wong-Foy, A. J. Matzger and M. S. Sanford, *J. Am. Chem. Soc.*, 2013, **135**, 10586; i) J. M. Taylor, K. W. Dawson and G. K. H. Shimizu, *J. Am. Chem. Soc.*, 2013, **135**, 1193; j) M. Yoon, K. Suh, S. Natarajan and K. Kim, *Angew. Chem., Int. Ed.*, 2013, **52**, 2688; k) Z. Xu, *Coord. Chem. Rev.*, 2006, **250**, 2745; l) O. M. Yaghi, G. M. Li and H. L. Li, *Nature*, 1995, **378**, 703; m) G. B. Gardner, D. Venkataraman, J. S. Moore and S. Lee, *Nature* 1995, **374**, 792; n) B. F. Hoskins and R. Robson, *J. Am. Chem. Soc.*, 1989, **111**, 5962.
- a) J. He, C. Yang, Z. Xu, M. Zeller, A. D. Hunter and J. Lin, *J. Solid State Chem.*, 2009, **182**, 1821; b) Z. Xu, *Metal-Organic Frameworks: Semiconducting Frameworks*, John Wiley & Sons, Ltd, Chichester, 2014.
- a) K.-K. Yee, N. Reimer, J. Liu, S.-Y. Cheng, S.-M. Yiu, J. Weber, N. Stock and Z. Xu, *J. Am. Chem. Soc.*, 2013, **135**, 7795; b) B. Li, Y. Zhang, D. Ma, Z. Shi and S. Ma, *Nat. Commun.*, 2014, **5**, 5537.
- a) L. Sun, T. Miyakai, S. Seki and M. Dincă, *J. Am. Chem. Soc.*, 2013, **135**, 8185; b) J. Cui and Z. Xu, *Chem. Commun.*, 2014, **50**, 3986; c) D. L.

- Turner, T. P. Vaid, P. W. Stephens, K. H. Stone, A. G. DiPasquale and A. L. Rheingold, *J. Am. Chem. Soc.*, 2008, **130**, 14.
5. a) S. Pullen, H. Fei, A. Orthaber, S. M. Cohen and S. Ott, *J. Am. Chem. Soc.*, 2013, **135**, 16997; b) H. Fei and S. M. Cohen, *J. Am. Chem. Soc.*, 2015, **137**, 2191; c) B. Gui, K.-K. Yee, Y.-L. Wong, S.-M. Yiu, M. Zeller, C. Wang and Z. Xu, *Chem. Commun.*, 2015, **51**, 6917.
6. W. J. Phang, H. Jo, W. R. Lee, J. H. Song, K. Yoo, B. S. Kim and C. S. Hong, *Angew. Chem., Int. Ed.*, 2015, Ahead of Print.
7. M. G. Goesten, J. Juan-Alcaniz, E. V. Ramos-Fernandez, K. B. S. S. Gupta, E. Stavitski, H. van Bekkum, J. Gascon and F. Kapteijn, *J. Catal.*, 2011, **281**, 177.
8. a) S. Biswas, J. Zhang, Z. Li, Y.-Y. Liu, M. Grzywa, L. Sun, D. Volkmer and P. Van Der Voort, *Dalton Trans.*, 2013, **42**, 4730; b) M. Lammert, S. Bernt, F. Vermoortele, D. E. De Vos and N. Stock, *Inorg. Chem.*, 2013, **52**, 8521; c) M. L. Foo, S. Horike, T. Fukushima, Y. Hijikata, Y. Kubota, M. Takata and S. Kitagawa, *Dalton Trans.*, 2012, **41**, 13791; d) B. Li, Y. Zhang, D. Ma, L. Li, G. Li, G. Li, Z. Shi and S. Feng, *Chem. Commun.*, 2012, **48**, 6151.
9. G. Chang, M. Huang, Y. Su, H. Xing, B. Su, Z. Zhang, Q. Yang, Y. Yang, Q. Ren, Z. Bao and B. Chen, *Chem. Commun.*, 2015, **51**, 2859.
10. a) D. F. Othmer, K. Kon and T. Igarashi, *Ind. Eng. Chem.*, 1948, **48**, 1258; b) V. H. Agreda, *Acetic Acid and its Derivatives*, CRC Press, 1992.
11. R. Jira, *Angew. Chem., Int. Ed.*, 2009, **48**, 9034.
12. a) S. Liang, J. Jasinski, G. B. Hammond and B. Xu, *Org. Lett.*, 2015, **17**, 162; b) J. Cordon, G. Jimenez-Oses, J. M. Lopez-de-Luzuriaga, M. Monge, M. E. Olmos and D. Pascual, *Organometallics*, 2014, **33**, 3823; c) L. Li, M. Zeng and S. B. Herzon, *Angew. Chem., Int. Ed.*, 2014, **53**, 7892; d) S. Wang, C. Miao, W. Wang, Z. Lei and W. Sun, *ChemCatChem*, 2014, **6**, 1612; e) W. Wang, A. Zheng, P. Zhao, C. Xia and F. Li, *ACS Catal.*, 2014, **4**, 321; f) W. E. Brenzovich, Jr., *Angew. Chem., Int. Ed.*, 2012, **51**, 8933; g) F.-X. Zhu, W. Wang and H.-X. Li, *J. Am. Chem. Soc.*, 2011, **133**, 11632.