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ARTICLE

Polydopamine-based synthesis of zeolite imidazolate framework ZIF-100 membrane with high H₂/CO₂ selectivity

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A highly permselective ZIF-100 molecular sieve membrane has been prepared on a polydopamine (PDA)-modified support. Attributing to the formation of strong covalent and non-covalent bonds between PDA and ZIF-100, the ZIF-100 nutrients are attracted and bound to the support surface, thus promoting the growth of well-intergrown and phase-pure ZIF-100 membranes. The developed ZIF-100 membranes show high H₂/CO₂ selectivity due to the outstanding CO₂ adsorption capacity of ZIF-100.

1. Introduction

The growing demand for energy and environmental issues has promoted the concept of “hydrogen economy”.¹ Currently, up to 90% of hydrogen is produced by steam-methane reforming (SMR) followed by the water-gas shift (WGS). Before it can be used in fuel cells, H₂ has to be purified from the SMR gas mixture which mainly contains CO₂.² Moreover, H₂-selective membranes are also desired in the pre-combustion technology of CO₂ sequestration.³ Compared with traditional separation methods like pressure swing adsorption (PSA) and cryogenic distillation, membrane-based separation techniques have attracted significant attention due to lower energy consumption and investment cost.⁴ In the recent 20 years, inorganic membranes such as zeolites membranes,⁵ Pd-based membranes,⁶⁻⁷ amorphous microporous silica membranes⁸ and carbon membranes⁹ have been developed for the separation of H₂ from CO₂.

Recent efforts are devoted to the fabrication of supported metal-organic framework (MOF) membranes due to their highly diversified structures and specific adsorption affinities.¹⁰⁻¹⁷ In particular, great attention has been paid to zeolite imidazolate framework (ZIF) membranes since 2009 because of their relative high stability and tunable pore size.¹⁸⁻¹⁹ So far, a series of ZIF membranes with small pore sizes, including ZIF-7 (0.30 nm),²⁰ ZIF-22 (0.30 nm),²¹ ZIF-8 (0.34 nm),²²⁻²⁴ ZIF-90 (0.35 nm)²⁵⁻²⁷ and ZIF-95 (0.37 nm),²⁸ were successively developed for gas separation. Despite much progress in the development of H₂-selective ZIF membranes, the development of thermally stable ZIF membranes with a high H₂/CO₂ selectivity is still desired.

Yaghi and co-workers have developed the novel ZIF-100 structure with the composition Zn₂₀(cbIM)₃₉(OH) through the reaction of Zn(O₃SCF₃)₂ with 5-chlorobenzimidazole (cbIM).²⁹ ZIF-100 was found to have a rather complex structure. The unit cell of ZIF-100 has a MOZ topology, which is constructed from 7524 atoms. This MOZ cage has a large inner sphere with a diameter of 35.6 Å and a constricted window aperture of only 3.35 Å. ZIF-100 shows a high affinity and capacity to CO₂, which results in an outstanding CO₂ uptake.²⁹ It is reported that ZIF-100 not only has high CO₂ capacity, but also a high thermal stability up to 500 °C.²⁹ It is worth to mention that the high CO₂ adsorption capacity as found for MOFs with accessible metal ions like MIL-53³⁰ or MOF-74³¹ is not common for ZIFs. However, the adsorption affinity of ZIF-100 for CO₂ outperforms ZIF-95 and BPL carbon³², which is widely used in industry for gas separation.

Although there have been a few studies of ZIF-100 powders on gas adsorption,³³ to date no ZIF-100 membrane with gas

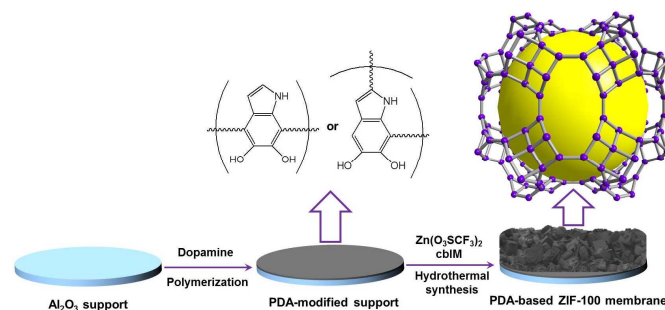


Fig. 1 Schematic diagram of the synthesis of ZIF-100 membranes on polydopamine-modified Al₂O₃ support.

separation performance has been reported. In this study, we report the synthesis of a dense ZIF-100 membrane by covalent modification of the substrate surface using polydopamine (PDA), as shown in Fig. 1. During this pre-modification process, dopamine can easily polymerize to PDA, and stick on different kinds of organic and inorganic material surfaces.³⁴ Then the PDA-modified surface can be used as a versatile platform for further reactions.³⁵⁻³⁷ Recently, PDA-based synthesis of ZIF-8 and zeolite LTA membranes has been developed to improve their gas separation performance and reproducibility.³⁸⁻⁴⁰ It can be expected, therefore, that ZIF-100 membranes prepared on PDA-functionalized supports will show higher gas separation performances and reproducibility of membrane preparation.

2. Experimental

2.1 Materials

Chemicals were used as received: zinc trifluoromethanesulfonate (98%, Aldrich), 5-chlorobenzimidazole (cbIM, 98%, Abcr), N, N-Dimethylformamide (DMF, water<50 ppm, Acros), dopamine hydrochloride (Sigma), tris(hydroxymethyl)amino-methane (Tris-HCl, ≥99.8%, Sigma-Aldrich). Porous α -Al₂O₃ disks (Fraunhofer Institute IKTS, former HITK/Inoceramic, Hermsdorf, Germany: 18 mm in diameter, 1.0 mm in thickness, 70 nm particles in the top layer) were used as supports.

2.2 Polydopamine (PDA) modification of the support surface

The PDA modification of the support surface was done according to a previous procedure.³⁸⁻⁴⁰ Dopamine hydrochloride (2 mg/mL) was dissolved in 10 mM Tris-HCl (pH 8.5) in an open to air vessel (diameter: 180 mm). And then the porous α -Al₂O₃ disks were placed face up into the dopamine solution (50 mL). The magnetic stirrer was placed in the middle of the vessel surrounded in a satellite-shape by 6 support discs. The solution was then stirred for 20 h at room temperature, leading to the polymerization of dopamine into PDA on the alumina support surface. The oxygen needed for the dopamine polymerization dissolved from air.

2.3 Synthesis of ZIF-100 membrane

The ZIF-100 membrane was prepared by a solvothermal reaction of trifluoromethanesulfonate and cbIM in DMF according to the previous report.²⁹ The PDA-modified or PDA-free α -Al₂O₃ supports were placed horizontally in a Teflon-lined stainless steel autoclave which was filled with 15 mL synthesis solution, and heated at 120 °C in air oven for 48 h. After solvothermal reaction, the ZIF-100 membranes were washed with DMF several times, and then dried in air at 110 °C over night.

2.4 Characterization of ZIF-100 membrane

The X-ray diffraction (XRD) patterns were recorded at room temperature under ambient conditions with X-ray diffractometer (D8 Advance, Bruker-AXS, with Cu-K α radiation at 40 kV and 40 mA). The morphology and thickness of the ZIF-100 membranes were characterized by scanning electron microscopy (SEM) at 2 keV and 10 μ A by using a JEOL Jeol-JSM-6700F with a cold field emission gun. By using the same SEM microscope the chemical composition of the cross-section of ZIF-100 was characterized by energy-dispersive X-ray spectroscopy (EDXS) at 20 kV and 20 μ A. The chemical composition of the cross-section of ZIF-100 layer was characterized by energy-dispersive X-ray spectroscopy (EDXS) using the same SEM microscope at 20 kV and 20 μ A.

2.5 Simulation models and methods

Grand canonical Monte Carlo (GCMC) simulation method in RASPA package⁴¹ was used to calculate the H₂/CO₂ mixed gas adsorption in ZIF-100. In the GCMC ensemble, the chemical potential, the volume, and the temperature are kept fixed as in the adsorption experiments. The chemical potential was related to the system pressure by the Peng-Robinson equation of state. In this work, the structure of ZIF-100 was constructed from the experimental single-crystal X-ray diffraction data.²⁹ The standard 12-6 Lennard-Jones (LJ) potential was used to model the dispersive and repulsive interatomic interactions. The Lorentz-Berthelot mixing rules were employed to calculate gas/framework parameters. The LJ parameters for the ZIF-100 atoms were obtained from the Dreiding force field,⁴² and if not available in Dreiding, from the Universal Force Field.⁴³ The partial charges of the ZIF-100 atoms were estimated using the CHELPG method⁴⁴ and density functional theory (DFT) calculation with the B3LYP method⁴⁵⁻⁴⁶ and the LANL2DZ basis set. A united-atom model was used for CH₄ with the LJ parameters from the TraPPE force field.⁴⁷ N₂ was mimicked by two-site models with a bond length of 1.10 Å.⁴⁸ Partial charges and LJ parameters for CO₂ were taken from the TraPPE force field. The electrostatic interactions for adsorbent-adsorbate and adsorbate-adsorbate were calculated by the Ewald summation technique.⁴⁹

2.6 Permeation of single gas and separation of mixed gases

For the single and mixed gas permeation, every ZIF-100 membrane was sealed in a permeation cell with silicone O-rings. In our Wicke Kallenbach permeation cell, we worked with feed and sweep gases at 1 bar. A calibrated gas chromatograph (HP6890) was used to detect the gas concentrations. The fluxes of both the feed and sweep gases were controlled by mass flow controllers. The flow rate on the feed side was kept constant for each gas with 50 mL min⁻¹, while the flow rate on the permeate side was kept at 50 mL min⁻¹ as well. N₂ was usually used as sweep gas, except for the measurement of N₂ permeance where CH₄ was used as sweep gas instead.

The permeance P is obtained by division of the flux by the transmembrane pressure difference, and the separation factor

$a_{i,j}$ of a binary mixture permeation is defined as the quotient of the molar ratios of the components (i, j) in the permeate, divided by the quotient of the molar ratio of the components (i, j) in the retentate. Since less than 1 % of the feed gas can pass through the membrane, the retentate composition is de facto identical with the feed composition.

To prove the feasibility of the Wicke-Kallenbach technique, the membranes were also tested under pressure difference. In this case, the pressure on the feed side was 4 bar (2 bar partial pressure of the binary mixture), while the pressure on the permeate side was 1 bar.

The gas chromatograph (GC) was calibrated every week anew with standard gas mixtures. The accuracy of the GC analysis of our H₂/CO₂ mixture with TCD detection is about ± 5 vol. %. Before gas permeation, the ZIF-100 membranes were activated at 100 °C with a heating rate of 0.2 °C min⁻¹ by using an equimolar H₂/CO₂ mixture in the Wicke-Kallenbach permeation apparatus. All permeation data were collected in steady state of permeation after 12 h equilibration time.

3. Results and discussion

3.1 Simulation study of gas adsorption isotherms of ZIF-100

The mixed gas adsorption isotherms in ZIF-100 were examined by molecular simulation. The adsorption of an equimolar H₂/CO₂ mixture in the pore structure of ZIF-100 cage at 25 °C and 1 bar is shown as a snapshot in Fig. 2. It can be seen clearly that a huge number of CO₂ molecules are adsorbed in the pores of ZIF-100, while only a few H₂ molecules are found. The simulation results of the gas adsorption isotherms of H₂, CO₂, CH₄ and N₂ at 25 °C are compared to the available experimental data²⁹ in the pressure range from 10 to 750 Torr in Fig. 3. As shown in Fig. 3, the simulated isotherms of different gases coincide well with the measured ones in the whole pressure range. The adsorption of CO₂ exceeds by far the adsorption of the other gases. The simulated adsorption isotherms of an equimolar H₂/CO₂

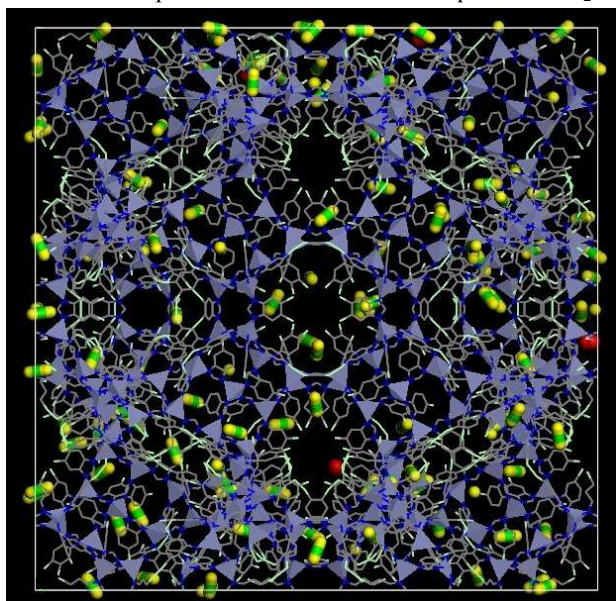


Fig. 2 Snapshot of adsorption of equimolar H₂/CO₂ mixture in ZIF-100 pores at 25 °C and 1 bar with CO₂ in yellow-green and H₂ in red.

mixture at 25 °C are shown in Fig. 4 as a function of the total pressure. As shown in Fig. 4, CO₂ is predominantly adsorbed over H₂ in ZIF-100 in the whole pressure range due to the much

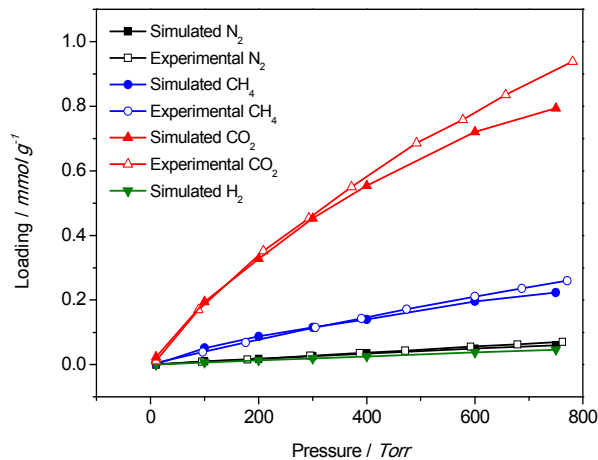


Fig. 3 Simulated single component adsorption isotherms of H₂, CO₂, CH₄ and N₂ at 25 °C in the pressure range from 10 to 750 Torr compared with the measured data²⁹.

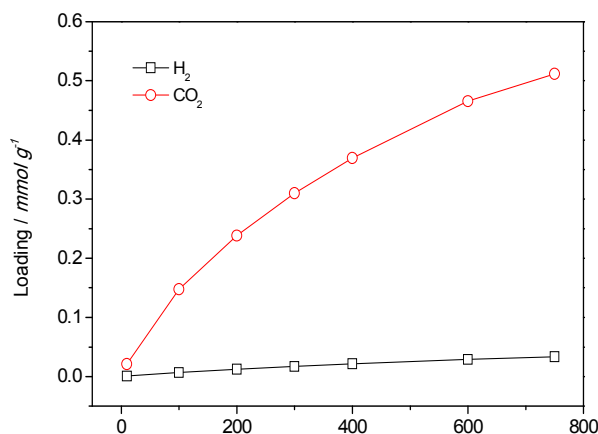


Fig. 4 Simulated mixed gas adsorption isotherms of equimolar H₂/CO₂ mixtures at 25 °C in the pressure range from 10 to 750 Torr. The pressure on the abscissa indicates the total pressure of the equimolar binary mixture.

stronger interaction between CO₂ and ZIF-100. As a result, the simulation study of the gas adsorption isotherm in ZIF-100 is in good accordance with the experimental data, and the experimentally found strong adsorption affinity of ZIF-100 to CO₂ was confirmed by the simulation.

3.2 Synthesis of ZIF-100 membrane

We first tried to grow a ZIF-100 layer simply by in-situ hydrothermal synthesis on the unmodified α -alumina support, as shown in Fig. 5 (a). However, it turned out extremely difficult to form a continuous ZIF-100 layer direct on the support due to the poor heterogeneous nucleation of ZIF-100 crystals on the α -alumina support surface. Our recent study indicated that pre-modification of the support surface with PDA can be employed³⁸ as a simple and powerful strategy to improve the nucleation of MOFs and zeolites on the alumina

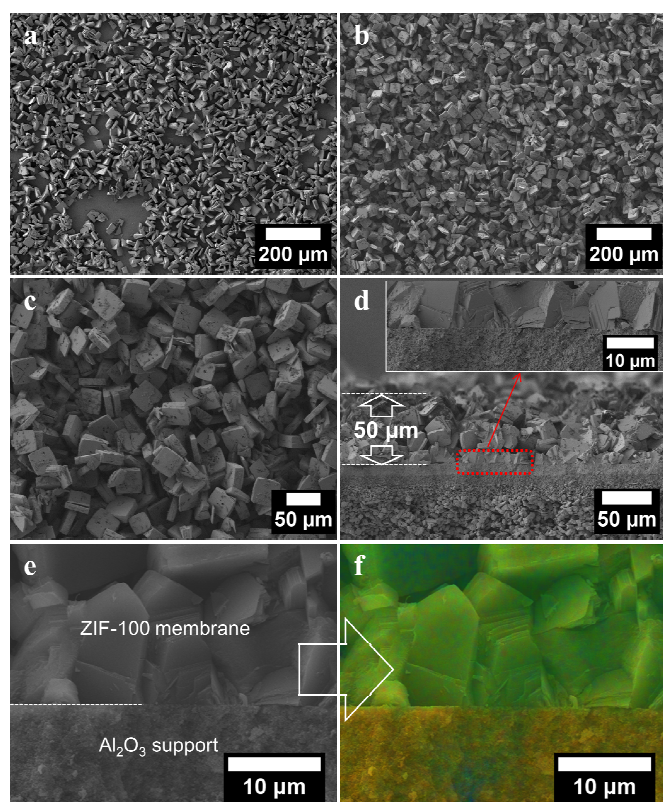


Fig. 5 SEM top views of (a) ZIF-100 membranes prepared on a PDA-free Al_2O_3 support and (b) (c) on a PDA-modified support, as well as (d) (e) the cross-section view of the ZIF-100 membrane with PDA-modification. The inset of (d) shows a magnified part of the ZIF-100 membrane at the interface support-ZIF-100 layer. (f) shows the EDXS mapping of the magnified cross section (green: Zn and Cl as tracers for ZIF-100; orange: Al as tracer for the support). The $\text{K}_{\alpha 1}$ transitions have been used.

support surface through the formation of strong covalent bonds with the MOF or zeolite crystals.³⁸⁻⁴⁰ In the present work, therefore, we tried to prepare a dense ZIF-100 membrane on the pre-modified support with PDA. As shown in Fig. 5 (b) and (c), a continuous ZIF-100 layer has formed on the PDA-modified supports. From the cross-section view shown in Fig. 5 (d) it follows, that big crystals are loosely packed on the top of the ZIF-100 membranes due to sedimentation, but as we can see from the magnified part (inset in Fig. 5d), a thin and well-intergrown ZIF-100 zone is formed at the interface to the support, thus controlling the selectivity for gas separation. Fig. 5(f) shows the EDXS mapping of the cross section of the ZIF-100 membrane near the support surface (Fig. 5e). A sharp transition can be seen between the ZIF-100 membrane (Zn and Cl signal) and the alumina support (Al signal) which means that there was no infiltration of the synthesis solution into the support. Also the above mentioned well-intergrown ZIF-100 zone near to the support can be seen in Fig. 5 (f).

The XRD patterns of ZIF-100 membranes are compared with ZIF-100 powder in Fig. 6. The results show that all diffraction peaks of the ZIF-100 layer prepared on either PDA-modified or bare supports match well with those of the ZIF-100 powder besides the $\alpha\text{-Al}_2\text{O}_3$ signals from the support, which indicates

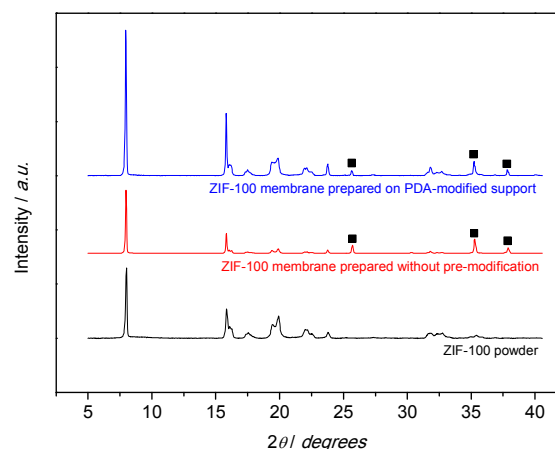


Fig. 6 XRD patterns of ZIF-100 membranes prepared on PDA-free and PDA-modified support, compared with the XRD patterns of ZIF-100 powder. (■) for Al_2O_3 support. Signals of ZIF-100 crystals are not marked.

that the ZIF-100 membrane prepared on the PDA-modified substrate exhibits the pure ZIF-100 structure. Moreover, the XRD pattern of PDA pre-modified ZIF-100 membrane shows higher ZIF-100 peak intensity relative to the membrane grown on bare alumina support, since the ZIF-100 membrane grown on PDA-based support is much denser and thicker.

3.3 Single gas permeation and mixture gas separation

It is worth to note that before the single gas permeation, the gas permeation performance of the mixed H_2/CO_2 gas pair on the PDA-modified alumina support was first measured and both the H_2 permeance and H_2/CO_2 separation factor kept unchanged when compared to the permeance through the bare Al_2O_3 support, which indicates that the PDA-layer itself is not gas-selective.

The single gas permeances of H_2 , CO_2 , N_2 and CH_4 through the ZIF-100 membrane prepared on PDA-modified support at

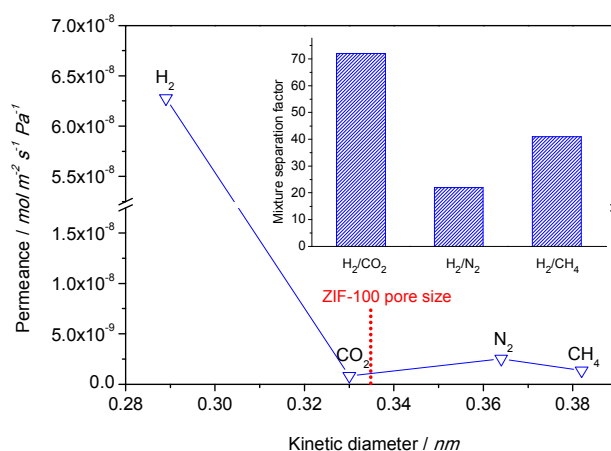


Fig. 7 Single gas permeances through the ZIF-100 membrane prepared by PDA-modification at 25 °C and 1 bar as a function of the kinetic diameter of permeated gases. The inset shows the mixture separation factors for H_2 over other gases from equimolar mixtures.

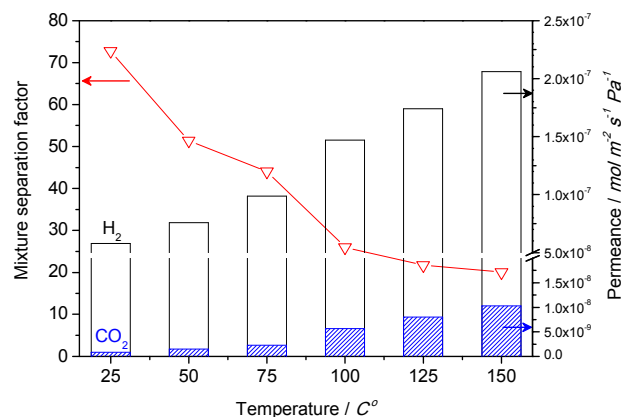


Fig. 8 Single gas permeances of H₂ and CO₂ and mixture separation factors for H₂/CO₂ from equimolar mixture on ZIF-100 membrane prepared by PDA-modification at 1 bar as a function of temperature.

25 °C and 1 bar as a function of the kinetic diameter of the gas molecules are shown in Fig. 7, and the results in details are summarized in Table 1. The inset gives the mixed separation factor for H₂ over other gases from their equimolar mixtures. It can be seen that the single gas permeance of H₂ at room temperature is $\sim 6.3 \times 10^{-8} \text{ mol}\cdot\text{m}^{-2}\cdot\text{s}^{-1}\cdot\text{Pa}^{-1}$, and the single gas permeances through the ZIF-100 membrane follow the order H₂>N₂>CH₄>CO₂, which leads to the highest H₂/CO₂ separation factor of 72. Although CO₂ has a smaller kinetic diameter than N₂ and CH₄, it was believed that the molecular sieving effect did not dominate the separation process, and the high H₂/CO₂ selectivity can be explained by diffusivity-solubility model of gas permeation instead. As reported previously,²⁹ only CO₂ can be retained in the pore structure of ZIF-100 while other small gas molecules will easily permeate through the framework since ZIF-100 shows a strong adsorption of CO₂ originating from the strong quadrupolar interactions of carbon with nitrogen atoms in the linkers of ZIF-100. The strong adsorption ability of CO₂ has been confirmed by the simulation study, as shown in section 3.1. The mobility of CO₂ was retarded by this strong adsorption, while H₂ can still pass through the pore network easily, leading to a high H₂/CO₂ selectivity. This experimental finding is similar to those reported on zeolite⁵⁰⁻⁵² and MOF membranes^{21,26,53-55}.

Fig. 8 shows the single gas permeance of H₂ and CO₂ as well

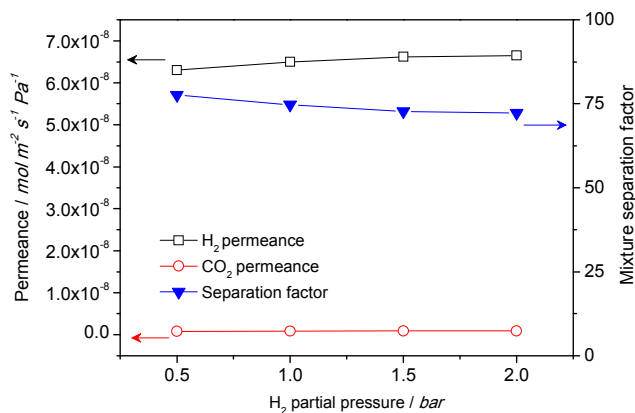


Fig. 9 H₂ and CO₂ permeances and H₂/CO₂ mixed gas selectivity of ZIF-100 membrane prepared by PDA-modification as a function of hydrogen partial pressure in the feed from equimolar mixtures at room temperature. Permeate pressure was kept constant at 1 bar.

as the mixture separation factors for equimolar H₂/CO₂ mixtures at 1 bar as a function of temperature from 25 to 150 °C. It can be seen that the H₂/CO₂ separation factors reduce gradually with increasing temperature in the temperature window of 25 to 100 °C, and then only slightly decrease when the temperature is higher than 100 °C. This trend can be explained by the interplay of adsorption and diffusion of H₂ and CO₂ in the pore structure of ZIF-100. At low temperatures, a large amount of CO₂ molecules is adsorbed in the ZIF-100 pores and only the highly mobile H₂ can diffuse through the membrane. As the temperature increases, less CO₂ becomes adsorbed and both H₂ and CO₂ can diffuse more easily in the resulting free volume, which leads to an enhancement of both H₂ and CO₂ permeance. Since the CO₂ permeance increases faster than the H₂ one, the H₂/CO₂ selectivity decreases with increasing temperature. After the temperature reached 100 °C, the effect of the CO₂ adsorption is substantially reduced and the gas separation behaviour of ZIF-100 becomes dominated by the kinetic molecular sieving effect since the width of the pore apertures of ZIF-100 is in the similar size to CO₂. However, because of the gate opening due to linker distortion, no sharp cut-off can be expected.⁵⁶⁻⁶³ Further, the ZIF-100 membrane shows completely reversible separation behaviour between 25 and 150 °C. The permeances measured during the cooling-down are well consistent with those during the heating-up.

Table 1 Single and mixed gas (with 1:1 binary mixtures) permeances as well as the ideal and mixture separation factors on ZIF-100 membrane prepared by pre-modification with PDA at 25 °C and 1 bar.

Gas _{i/j}	Knudsen constant	Separation performance of ZIF-100 membrane					
		Single gas			Mixed gases		
		Permeances (i) (mol·m ⁻² ·s ⁻¹ ·Pa ⁻¹)	Permeances (j) (mol·m ⁻² ·s ⁻¹ ·Pa ⁻¹)	Ideal separation factor	Permeances (i) (mol·m ⁻² ·s ⁻¹ ·Pa ⁻¹)	Permeances (j) (mol·m ⁻² ·s ⁻¹ ·Pa ⁻¹)	Mixture separation factor (α)
H ₂ /CO ₂	4.7	6.3 × 10 ⁻⁸	8.1 × 10 ⁻¹⁰	77	5.8 × 10 ⁻⁸	8.0 × 10 ⁻¹⁰	72
H ₂ /N ₂	3.7		2.5 × 10 ⁻⁹	25	6.2 × 10 ⁻⁸	2.7 × 10 ⁻⁹	22
H ₂ /CH ₄	2.8		1.4 × 10 ⁻⁹	46	5.8 × 10 ⁻⁸	1.4 × 10 ⁻⁹	41

Table 2 Single gas permeances of H₂ and CO₂ and mixture separation factors of H₂/CO₂ from equimolar mixtures at room temperature and 1 bar feed/1 bar permeate of 5 tested ZIF-100 membranes showing the reproducibility of membrane preparation and testing.

	H ₂ permeance (mol·m ⁻² ·s ⁻¹ ·Pa ⁻¹)	CO ₂ permeance (mol·m ⁻² ·s ⁻¹ ·Pa ⁻¹)	Mixture separation factor H ₂ /CO ₂
1	6.3 × 10 ⁻⁸	8.1 × 10 ⁻¹⁰	72
2	6.5 × 10 ⁻⁸	8.0 × 10 ⁻¹⁰	76
3	6.1 × 10 ⁻⁸	7.9 × 10 ⁻¹⁰	68
4	6.6 × 10 ⁻⁸	8.1 × 10 ⁻¹⁰	75
5	5.8 × 10 ⁻⁸	7.8 × 10 ⁻¹⁰	70

In addition, the ZIF-100 membrane can keep its high H₂/CO₂ selectivity when the feed pressure increases. As shown in Fig. 9, when the H₂ and CO₂ partial pressure increased from 0.5 to 2 bar (corresponding to a change in the total feed pressure from 1 to 4 bar), the H₂/CO₂ separation factor reduces only slightly which is a direct proof that the membrane does not contain macroscopic defects.

Moreover, it was also found that the synthesis method with pre-modification of the α -alumina supports by PDA can contribute to a higher reproducibility of ZIF-100 membrane preparation, as shown in Table 2. The gas separation performances of 5 different membranes prepared with the same method were tested and the mixture separation factors of H₂/CO₂ do not scatter more than $\pm 10\%$.

4. Conclusions

Dense and phase-pure ZIF-100 membranes have been prepared on PDA-modified alumina support at 120 °C for 48 h. After the pre-modification with PDA, the ZIF-100 nutrients were attached to the support surface through the formation of covalent chemical bonds. The mixture separation factors of H₂/CO₂, H₂/N₂ and H₂/CH₄ through the PDA-based ZIF-100 membrane were 72, 22 and 41 at room temperature and 1 bar. The high H₂/CO₂ selectivity is ascribed to the high CO₂ uptake behaviour of ZIF-100 and the small window aperture of 3.35 Å.

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Notes and references

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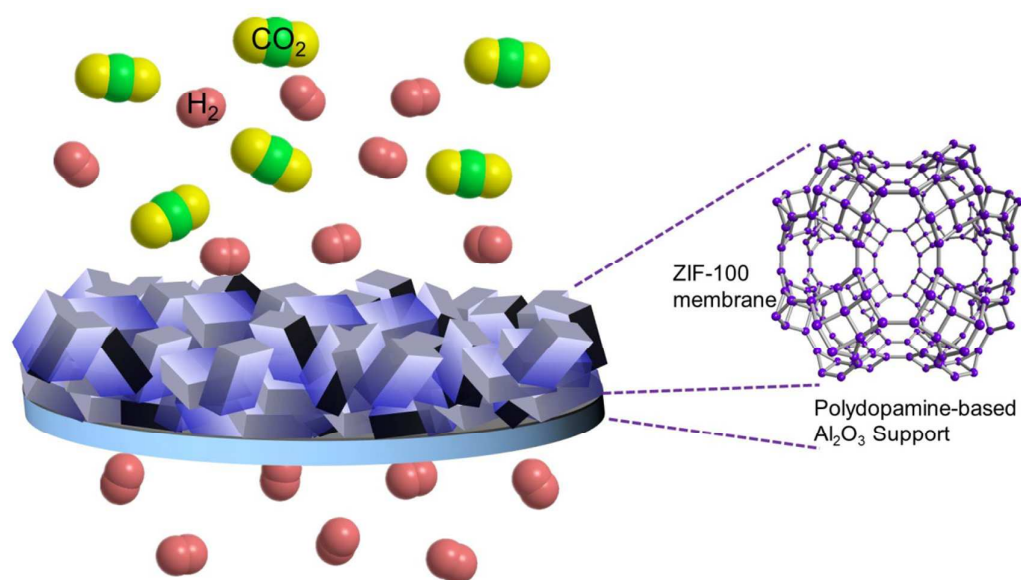
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- S. Dunn, *Int. J. Hydrogen Energy*, 2002, **27**, 235.
- J. R. Rostrup-Nielsen and T. Rostrup-Nielsen, *CATTECH*, 2002, **6**, 150.
- A. A. Olagire, *Energy*, 2010, **35**, 2610.
- A. J. Brown, N. A. Brunelli, K. Eum, F. Rashidi, J. R. Johnson, W. J. Koros, C. W. Jones and S. Nair, *Science*, 2014, **345**, 72.
- M. Hong, S. Li, J. L. Falconer and R. D. Noble, *J. Membr. Sci.*, 2008, **307**, 277.
- F. Gallucci, E. Fernandez, P. Corengia and M. S. Annaland, *Chem. Eng. Sci.*, 2013, **92**, 40.
- S. Uemiyama, T. Matsuda, E. Kikuchi, *J. Membr. Sci.*, 1991, **56**, 315.
- R. M. de Vos, H. Verweij, *Science*, 1998, **279**, 1710.
- M. B. Shiflett, H. C. Foley, *Science*, 1999, **285**, 1902.
- H. Li, M. Eddaoudi, M. O. Keffe and O. M. Yaghi, *Nature*, 1999, **402**, 276.
- M. Dincă, A. F. Yu and J. R. Long, *J. Am. Chem. Soc.*, 2006, **128**, 8904.
- S. Hermes, F. Schroder, R. Chelmoski, C. Woll and R. A. Fischer, *J. Am. Chem. Soc.*, 2005, **127**, 13744.
- R. Ranjan and M. Tsapatsis, *Chem. Mater.*, 2009, **21**, 4920.
- A. Bétard, H. Bux, S. Henke, D. Zacher, J. Caro and R. A. Fischer, *Micropor. Mesopor. Mater.*, 2012, **150**, 76.
- Y. Yoo, Z. Lai and H. K. Jeong, *Micropor. Mesopor. Mater.*, 2009, **123**, 100.
- T. Rodenas, M. van Dalen, E. Garcia-Pérez, P. Serra-Crespo, B. Zornoza, F. Kapteijn and J. Gascon, *Adv. Funct. Mater.*, 2014, **24**, 249.
- F. Cacho-Bailo, B. Seoane, C. Téllez, J. Coronas, *J. Membr. Sci.*, 2014, **464**, 119.
- R. Banerjee, A. Phan, B. Wang, C. Knobler, H. Furukawa, M. O'Keeffe and O. M. Yaghi, *Science*, 2008, **319**, 939.
- A. Phan, C. J. Doonan, F. J. Uribe-romo, C. B. Knobler, M. O'Keeffe and O. M. Yaghi, *Acc. Chem. Res.*, 2010, **43**, 58.
- Y. Li, F. Liang, H. Bux, A. Feldhoff, W. Yang and J. Caro, *Angew. Chem.*, 2010, **122**, 558; *Angew. Chem. Int. Ed.*, 2010, **49**, 548.
- A. Huang, H. Bux, F. Steinbach and J. Caro, *Angew. Chem. Int. Ed.*, 2010, **49**, 4958.
- H. Bux, F. Liang, Y. Li, J. Cravillon, M. Wiebcke, J. Caro, *J. Am. Chem. Soc.*, 2009, **131**, 16000.
- H. Bux, A. Feldhoff, J. Cravillon, M. Wiebcke, Y. Li and J. Caro, *Chem. Mater.*, 2011, **23**, 2262.
- Y. Pan and Z. Lai, *Chem. Commun.*, 2011, **47**, 10275.
- A. Huang, W. Dou and J. Caro, *J. Am. Chem. Soc.*, 2010, **132**, 15562.
- A. Huang and J. Caro, *Angew. Chem. Int. Ed.*, 2011, **50**, 4979.
- A. Huang, N. Wang, C. Kong and J. Caro, *Angew. Chem. Int. Ed.*, 2012, **51**, 10551.
- A. Huang, Y. Chen, N. Wang, Z. Hu, J. Jiang and J. Caro, *Chem. Commun.*, 2012, **48**, 10981.
- B. Wang, A. P. Côté, H. Furukawa, M. O'Keeffe and O. M. Yaghi, *Nature*, 2008, **453**, 207.
- S. Bourrelly, P. L. Llewellyn, C. Serre, F. Millange, T. Loiseau and G. Férey, *J. Am. Chem. Soc.* 2009, **131**, 6326.
- L. Valenzano, B. Civalieri, S. Chavan, G. T. Palomino, C. O. Areán, S. Bordiga, *J. Phys. Chem. C*, 2010, **114**, 11185.
- S. Sircar, T. C. Golden and M. B. Rao, *Carbon*, 1996, **34**, 1.

- 33 M. Prakash, N. Sakhavand and R. Shahsavari, *J. Phys. Chem. C.*, 2013, **117**, 24407.
- 34 H. Lee, S. M. Dellatore, W. M. Miller and P. B. Messersmith, *Science*, 2007, **318**, 426.
- 35 H. Lee, J. Rho and P. B. Messersmith, *Adv. Mater.*, 2009, **21**, 431.
- 36 J. Ryu, S. H. Ku, H. Lee and C. B. Park, *Adv. Funct. Mater.*, 2010, **20**, 2132.
- 37 D. Ling, W. Park, Y. I. Park, N. Lee, F. Li, C. Song, S. Yang, S. H. Choi, K. Na and T. Hyeon, *Angew. Chem., Int. Ed.*, 2011, **50**, 11360.
- 38 Q. Liu, N. Wang, J. Caro and A. Huang, *J. Am. Chem. Soc.*, 2013, **135**, 17679.
- 39 A. Huang, Q. Liu, N. Wang and J. Caro, *J. Mater. Chem. A.*, 2014, **2**, 8246.
- 40 C. Yuan, Q. Liu, H. Chen and A. Huang, *RSC Adv.*, 2014, **4**, 41982.
- 41 D. Dubbeldam, S. Calero, D. E. Ellis and R. Q. Snurr, eds. *RASPA 1.0. Northwestern University: Evanston, IL* 2008.
- 42 S. L. Mayo, B. D. Olafson and W. A. Goddard, *J. Phys. Chem.*, 1990, **94(26)**, 8897.
- 43 A. K. Rappe, C. J. Casewit, K. S. Colwell, W. A. Goddard III, W. M. Skiff, *J. Am. Chem. Soc.*, 1992, **114(25)**, 10024.
- 44 C. M. Breneman and K. B. Wiberg, *J. Comput. Chem.*, 1990, **11(3)**, 361.
- 45 C. Lee, W. Yang and R. G. Parr, *Phys. Rev. B.*, 1988, **37(2)**, 785.
- 46 A. D. Becke, *J. Chem. Phys.*, 1993, **98**, 5648.
- 47 M. G. Martin and J. I. Siepmann, *J. Phys. Chem. B.*, 1998, **102(14)**, 2569.
- 48 C. S. Murthy, K. Singer, M. L. Klein and I. R. McDonald, *Mol. Phys.*, 1980, **41(6)**, 1387.
- 49 P. P. Ewald, *Ann. Phys-New York*, 1921, **64(3)**, 253.
- 50 A. Huang, N. Wang and J. Caro, *J. Membr. Sci.*, 2012, **389**, 272.
- 51 A. Huang and J. Caro, *J. Mater. Chem.*, 2011, **21**, 11424.
- 52 N. Wang, Y. Liu, A. Huang and J. Caro, *Micropor. Mesopor. Mater.*, 2014, **192**, 8.
- 53 Y. Li, F. Liang, H. Bux, A. Feldhoff, W. Yang and J. Caro, *Angew. Chem. Int. Ed.*, 2010, **49**, 548.
- 54 A. Huang, Y. Chen, N. Wang, Z. Hu, J. Jiang and J. Caro, *Chem. Commun.*, 2012, **48**, 10981.
- 55 N. Wang, A. Mundstock, Y. Liu, A. Huang and J. Caro, *Chem. Eng. Sci.*, doi:10.1016/j.ces.2014.10.037.
- 56 C. Gücüyener, J. van den Bergh, J. Gascon and F. Kapteijn, *J. Am. Chem. Soc.*, 2010, **132**, 17704.
- 57 D. Fairen-Jimenez, S. A. Moggach, M. T. Wharmby, P. A. Wright, S. Parsons and T. Düren, *J. Am. Chem. Soc.*, 2001, **133**, 8900.
- 58 S. A. Moggach, T. D. Bennett and A. K. Cheetham, *Angew. Chem.*, 2009, **121**, 7221.
- 59 T. Chokbunpiam, R. Chanajaree, T. Remsungnen, O. Saengsawang, S. Fritzsche, C. Chmelik, J. Caro, W. Janke and S. Hannongbua, *Micropor. Mesopor. Mater.*, 2014, **187**, 1.
- 60 S. Aguado, G. Bergeret, M. P. Titus, V. Moizan, C. Nieto-Draghi, N. Bats and D. Farrusseng, *New J. Chem.*, 2011, **35**, 546.
- 61 J. van den Bergh, C. Gücüyener, E. A. Pidko, E. J. M. Hensen, J. Gascon and F. Kapteijn, *Chem. Eur. J.*, 2011, **17**, 8832.
- 62 D. L. Chen, N. Wang, F. F. Wang, J. Xie, Y. Zhong, W. Zhu, J. K. Johnson and R. Krishna, *J. Phys. Chem. C*, 2014, **118**, 17831.
- 63 C. O. Ania, E. García-Pérez, M. Haro, J. J. Gutiérrez-Sevillano, T. Valdés-Solís, J. B. Parra and S. Calero, *J. Phys. Chem. Lett.*, 2012, **3**, 1159.

Table of content graphics



A novel ZIF-100 membrane with high H_2/CO_2 selectivity has been developed on polydopamine-based alumina support.