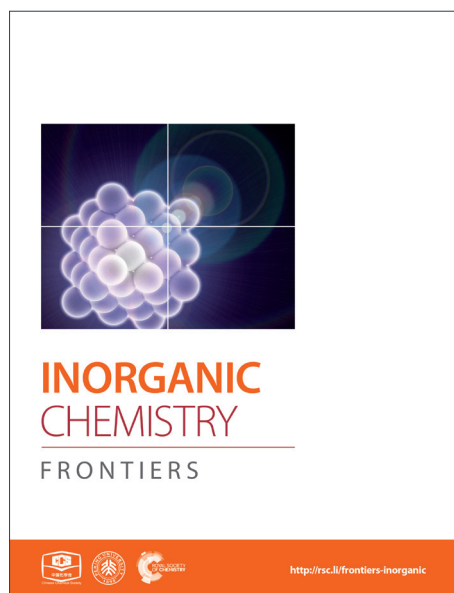
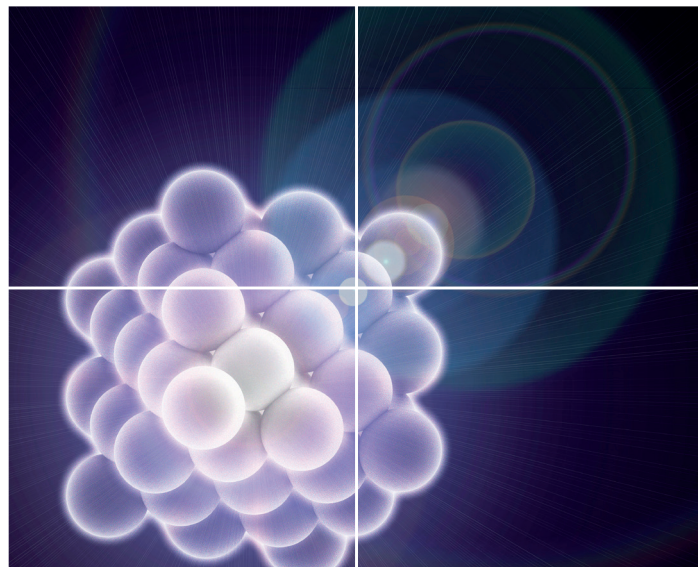


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Explosive Werner-type cobalt(III) complexes[†]

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Dedicated to Professor Konrad Seppelt on the occasion of his 70th birthday

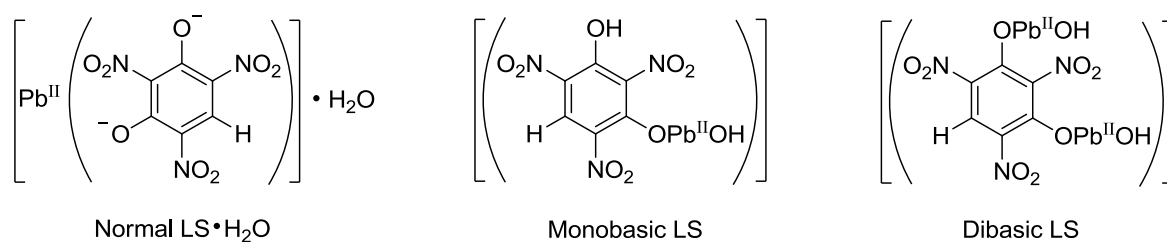
[†] Electronic supplementary information (ESI) available: ORTEP drawings and complete X-ray crystallographic data for **2**, **3**, **4**, **5**, **7**, **9**, **10**, **12**, and **13** in PDF and CIF format. CCDC 993736-993744. For crystallographic data in CIF or other electronic format see DOI:

A series of potentially explosive Werner-type cobalt(III) complexes comprising the anions azotetrazolate, nitrotetrazolate, picrate and dipicrylamide have been prepared *via* simple metathetical routes. Treatment of $[\text{Co}(\text{NH}_3)_5\text{NO}_2]\text{Cl}_2$, *trans*- $[\text{Co}(\text{NH}_3)_4(\text{py})\text{NO}_2]\text{Cl}_2$ (py = pyridine), *trans*- $[\text{Co}(\text{NH}_3)_4(\text{NO}_2)_2]\text{Cl}$, and $[\text{Co}(\text{NH}_3)_5\text{N}_3]\text{Cl}_2$ with equimolar amounts of disodium azotetrazolate, $(\text{Na}_2\text{C}_2\text{N}_{10}\cdot 5\text{H}_2\text{O})$, **1**, in aqueous solutions afforded the new cobalt(III) azotetrazolate salts $[\text{Co}(\text{NH}_3)_5\text{NO}_2](\text{C}_2\text{N}_{10})\cdot 2\text{H}_2\text{O}$ (**2**), *trans*- $[\text{Co}(\text{NH}_3)_4(\text{py})\text{NO}_2](\text{C}_2\text{N}_{10})\cdot 2\text{H}_2\text{O}$ (**3**), *trans*- $[\text{Co}(\text{NH}_3)_4(\text{NO}_2)_2](\text{C}_2\text{N}_{10})$ (**4**), and $[\text{Co}(\text{NH}_3)_5\text{N}_3](\text{C}_2\text{N}_{10})\cdot \text{H}_2\text{O}$ (**5**) in moderate to excellent yields (46-88%). Similar treatment of *trans*- $[\text{Co}(\text{NH}_3)_4(\text{NO}_2)_2]\text{Cl}$ with 1 equiv. of sodium 5-nitrotetrazolate dihydrate (= NaNT, **6**) afforded the novel cobalt(III) 5-nitrotetrazolate derivative *trans*- $[\text{Co}(\text{NH}_3)_4(\text{NO}_2)_2](\text{NT})\cdot \text{H}_2\text{O}$ (**7**) as orange, rectangular prismatic crystals in 64% yield. Two complex cobalt(III) picrates, *trans*- $[\text{Co}(\text{NH}_3)_4(\text{NO}_2)_2](\text{picrate})\cdot \text{H}_2\text{O}$ (**9**) and $[\text{Co}(\text{NH}_3)_5\text{N}_3](\text{picrate})_2$ (**10**), were prepared in a similar manner from the corresponding chloride precursors and equimolar amounts of sodium picrate. The reaction of *trans*- $[\text{Co}(\text{NH}_3)_4(\text{NO}_2)_2]\text{Cl}$ with sodium dipicrylamide (= NaDPA) in a 1:1 molar ratio gave the first cobalt(III) dipicrylamide, *trans*- $[\text{Co}(\text{NH}_3)_4(\text{NO}_2)_2](\text{DPA})\cdot \text{H}_2\text{O}$ (**12**). Finally, the highly explosive, dark blue-green dichroitic non-electrolyte complex *mer*- $[\text{Co}(\text{en})(\text{py})(\text{N}_3)_3]$ (**13**) was formed upon treatment of $[\text{Co}(\text{en})(\text{py})_2(\text{NH}_3)\text{Cl}]\text{Cl}_2\cdot \text{H}_2\text{O}$ with excess NaN_3 in hot water (93% yield). The molecular and crystal structures of **2**, **3**, **4**, **5**, **7**, **9**, **10**, **12**, and **13** have been determined by single-crystal X-ray diffraction. In the solid state, all compounds comprise extensive hydrogen-bonded supramolecular networks. Representative studies of the energetic properties (impact and friction sensitivity, combustion) revealed that some of the new compounds can be classified as primary explosives.

INTRODUCTION

Primary explosives are highly sensitive explosive compounds which are used to initiate large amounts of secondary explosives such as 2,4,6-trinitrotoluene (= TNT)¹ in initiating devices like primers and detonators for commercial and military applications. Historically, one of the first practical primary explosives was mercury fulminate, $\text{Hg}(\text{CNO})_2$ ("Knallquecksilber"), which has been widely used for almost 100 years.² Later, mercury fulminate was replaced by lead(II) azide, $\text{Pb}(\text{N}_3)_2$,³ and several forms of lead

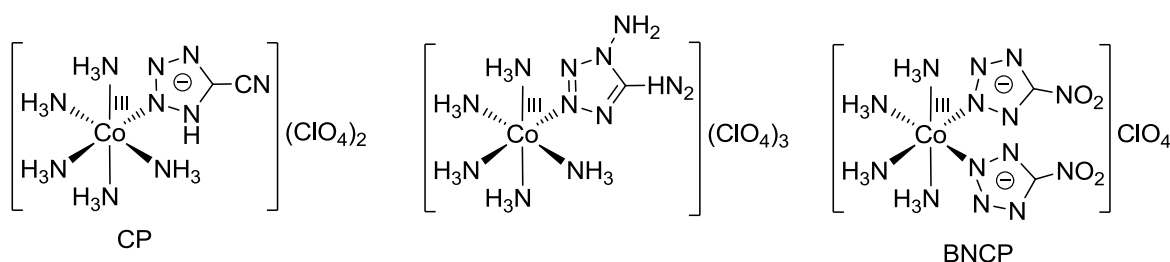
styphnate (= LS, Scheme 1)⁴ which are being generally used as primary explosive until today.



Scheme 1

However, lead(II) azide inherits the severe disadvantages that highly toxic hydrozoic acid, HN₃, can be formed under certain conditions,³ and the use of all these materials is always associated with lead pollution of the environment.⁵ For this reason, the search for environmentally acceptable ("green") primary explosives is currently a hot topic in this field.⁶ One of the most useful approaches is the development of nitrogen-rich energetic compounds based on tetrazole and tetrazine derivatives.⁷ A highly promising candidate to come out of these research efforts is the recently reported copper(I) 5-nitrotetrazolate (= DBX-1).⁸ Easily prepared DBX-1 has been shown to be one of the best "drop-in" lead(II) azide replacements to date due to its high thermal stability and comparable safety and performance properties.

Several Werner-type cobalt(III) complexes comprising 5-nitrotetrazolate and related ligands have also been synthesized and reported to have primary explosive properties. Especially the complexes shown in Scheme 2, pentaammine(5-cyanotetrazolato-*N*²)cobalt(III) perchlorate (CP)⁹, pentaammine(4,5-diaminotetrazole-*N*²)cobalt perchlorate,¹⁰ and tetraammine-bis(5-nitrotetrazolato-*N*²)cobalt(III) perchlorate (BNCP),¹¹ were shown to have suitable energetic properties. However, their perchlorate content makes them unacceptable as alternative primary explosives because perchlorate has various adverse health effects.¹²



Scheme 2

Other potentially explosive Werner-type cobalt(III) complexes have frequently been described in the literature without the aim of any practical uses. These comprise mainly some long-known azido complexes. While several series of complexes such as $[\text{Co}(\text{NH}_3)_5\text{N}_3]^{2+}$,¹³ *cis*- and *trans*- $[\text{Co}(\text{NH}_3)_4(\text{N}_3)_2]^+$,¹⁴ and *cis*- and *trans*- $[\text{Co}(\text{en})_2(\text{N}_3)_2]^+$ (en = ethylene-1,2-diamine)¹⁵ are often thermally quite stable (with the exception of some azides and perchlorates), there are several cases of highly explosive species. For example, several salts of the $[\text{Co}(\text{N}_3)_6]^{3-}$ anion were found to be quite sensitive to impact and friction,¹⁶ and the compound *trans*-tetraammine-diazido-cobalt(III)-*trans*-diammine-tetraazido-cobaltate(III), $[\text{Co}(\text{NH}_3)_4(\text{N}_3)_2][\text{Co}(\text{NH}_3)_2(\text{N}_3)_4]$, has been described as a "dangerous detonator", but its energetic properties have not been studied in detail.¹⁷ Also notable is the (in)famous green non-electrolyte complex *mer*- $[\text{Co}(\text{NH}_3)_3(\text{N}_3)_3]$. It has been reported that this compound is highly impact-sensitive and can even explode upon grinding under water.¹⁸

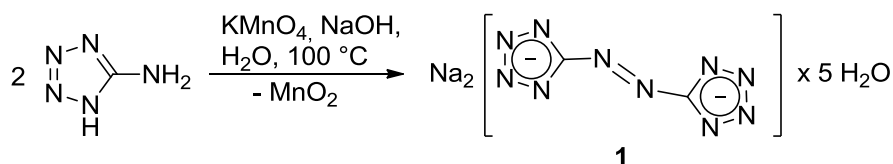
In this contribution we report the synthesis, structural characterization, and energetic properties of a series of new explosive Werner-type cobalt(III) complexes. Our synthetic protocol involves combination of complex cobalt(III) cations with various nitrogen-rich or oxygen-rich, sensitive, secondary high explosive anions. For the present study, azotetrazolate, 5-nitrotetrazolate, picrate, and dipicrylamide were chosen as suitable representative energetic anions. Unlike the cobalt(III) complexes depicted in Scheme 2, the new compounds reported here contain the energetic anions in the outer sphere and not directly coordinated to the central Co^{3+} cation.

RESULTS AND DISCUSSION

Azotetrazolates

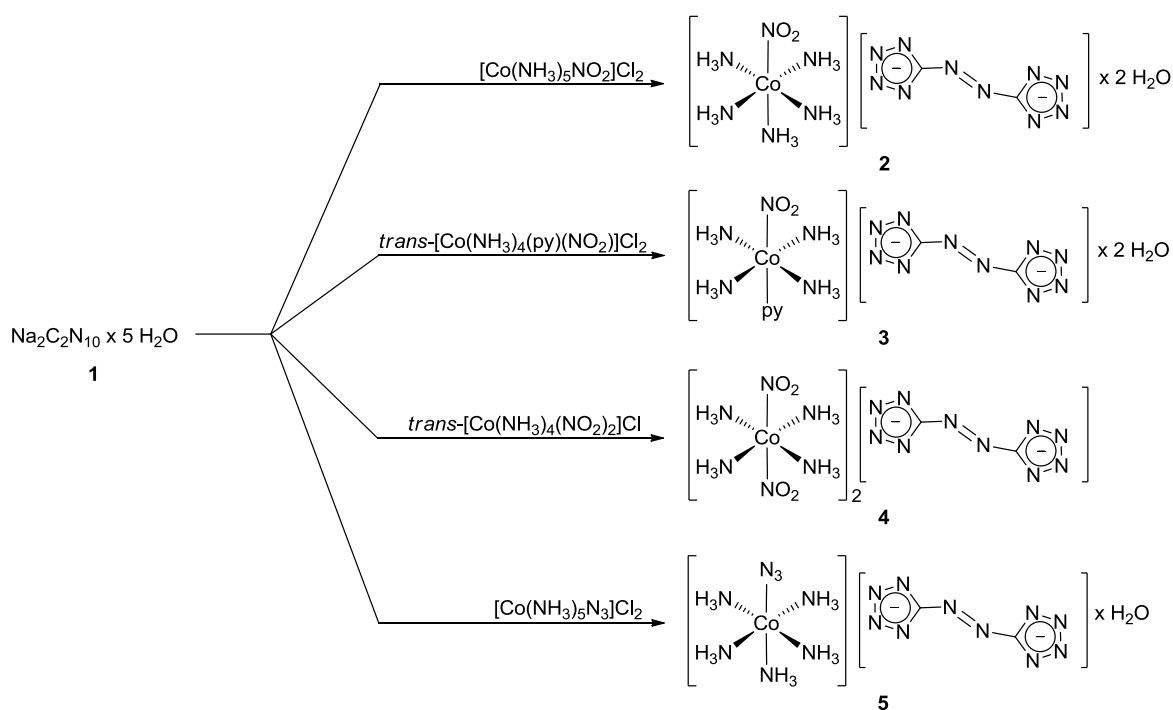
The azotetrazolate dianion, $(\text{C}_2\text{N}_{10}^{2-})$, has been found to be extremely useful in the design of new nitrogen-rich energetic salts.¹⁹ The disodium salt $\text{Na}_2\text{C}_2\text{N}_{10}\cdot 5\text{H}_2\text{O}$ (**1**) is the most

readily accessible azotetrazolate precursor described in the literature. It was first prepared by Thiele more than 100 years ago by oxidation of aminotetrazole with KMnO_4 in boiling sodium hydroxide solution (Scheme 3). This rather harsh method provides disodium azotetrazolate as its pentahydrate in the form of large, bright yellow crystals.²⁰



Scheme 3

New Werner-type cobalt(III) complexes containing the azotetrazolate anion were synthesized according to Scheme 4 by metathetical reactions of chloride precursors with equimolar amounts of **1**. As starting materials, several easily accessible penta- and tetraammincobalt(III) chlorides containing nitro and azido ligands were chosen ($[\text{Co}(\text{NH}_3)_5\text{NO}_2]\text{Cl}_2$, *trans*- $[\text{Co}(\text{NH}_3)_4(\text{py})\text{NO}_2]\text{Cl}_2$ (py = pyridine), *trans*- $[\text{Co}(\text{NH}_3)_4(\text{NO}_2)_2]\text{Cl}$, and $[\text{Co}(\text{NH}_3)_5\text{N}_3]\text{Cl}_2$).



Scheme 4

All reactions illustrated in Scheme 4 were carried out at room temperature in concentrated aqueous solutions. The products $[\text{Co}(\text{NH}_3)_5\text{NO}_2](\text{C}_2\text{N}_{10})\cdot 2\text{H}_2\text{O}$ (**2**, 63% yield), *trans*- $[\text{Co}(\text{NH}_3)_4(\text{py})\text{NO}_2](\text{C}_2\text{N}_{10})\cdot 2\text{H}_2\text{O}$ (**3**, 46%), *trans*- $[\text{Co}(\text{NH}_3)_4(\text{NO}_2)_2](\text{C}_2\text{N}_{10})$ (**4**, 87%), and $[\text{Co}(\text{NH}_3)_5\text{N}_3](\text{C}_2\text{N}_{10})\cdot \text{H}_2\text{O}$ (**5**, 88%) crystallized within 24 h upon undisturbed standing of the reaction mixtures. The nitro complexes **2-4** formed yellow to orange crystals, whereas the azido complex **5** crystallized in the form of dark red rods. Recrystallization from hot water was possible but not necessary, as the products directly obtained from the original reaction mixtures were already quite pure. Routine characterization was mainly based on IR data and elemental analyses. As expected, these data provided only some basic information on the nature of the products. Thus they will be briefly discussed here only for compound $[\text{Co}(\text{NH}_3)_5\text{NO}_2](\text{C}_2\text{N}_{10})\cdot 2\text{H}_2\text{O}$ (**2**) as a typical example. Elemental analysis was consistent with the formation of the expected product as the dihydrate. The IR spectrum showed two intense absorptions which are characteristic for the free azotetrazolate dianion. A strong band around $730\text{-}740\text{ cm}^{-1}$ can be assigned to the asymmetrical C-N=N stretching mode of the azo group, whereas a strong band around $1390\text{-}1400\text{ cm}^{-1}$ is attributable to the asymmetrical N-C=N stretching mode of the ring. These characteristic bands are more or less the same in all compounds containing uncoordinated azotetrazolate dianions.¹⁹ In the IR spectrum of **2**, the two bands appear at 740 and 1402 cm^{-1} . While ^1H NMR data were unavailable for the new complexes **2-5**, the presence of uncoordinated azotetrazolate dianions could also be verified by the ^{13}C NMR data, measured in DMSO-*d*₆ solutions. In all four cases, the signal of the carbon atom of the azotetrazolate dianion appeared in the ^{13}C NMR spectra at $\delta = 173.1$, which is characteristic for salt-like azotetrazolates.¹⁹ In the ^{13}C NMR spectrum of **3** the signals of the coordinated pyridine could be observed at $\delta = 8.4, 7.8,$ and 7.3 .

All four new compounds were also structurally characterized by single-crystal X-ray diffraction. In all cases, suitable single-crystals were obtained directly from the original reaction mixtures. A summary of the crystallographic data and refinement parameters for all nine crystal structures reported in this study are given in Table 1. Figures 1-12 show photographs of the single-crystals, ORTEP drawings of the molecular structures including selected bond lengths and angles, as well as packing diagrams of the hydrogen-bonded networks in the solid state.

Table 1 here

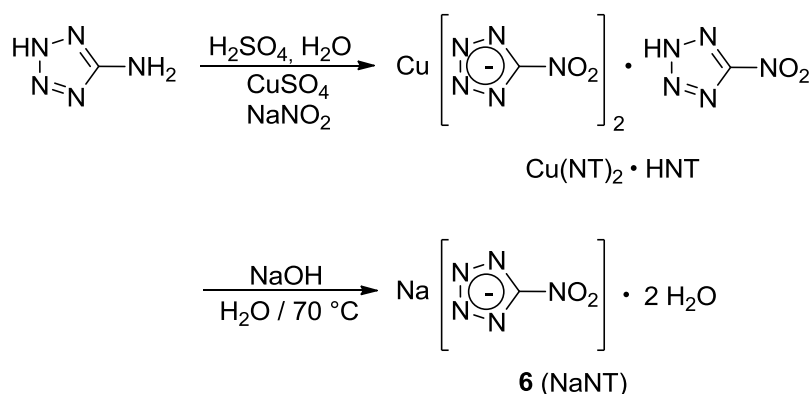
Figures 1-12 here

In the solid state, the compounds **2-5** consist of the complex cobalt(III) cations and azotetrazolate dianions in the respective stoichiometric ratio plus one (**5**) or two (**2,3**) equivalents of water of crystallization. Only the dinitro complex *trans*-[Co(NH₃)₄(NO₂)₂]₂(C₂N₁₀) (**4**) crystallizes water-free. The individual ions display no unusual structural features. The hexa-coordinated cobalt(III) cations show no significant deviation from the regular octahedral coordination geometry. Taken as a typical example, the *cis*-N-Co-N angles in the *trans*-[Co(NH₃)₄(NO₂)₂]⁺ cation in **4** are in the very narrow range of 87.86(8) - 91.10(7)°, whereas the *trans*-N-Co-N angles range from 177.01(6)° to 179.77(6)°. The central N=N double-bond length in the azotetrazolate dianion in **4** is 1.251(3) Å, whereas the N-N distances in the tetrazole rings clearly indicate delocalization of the negative charge because of bond homogenization with N-N distances of 1.339(2) Å (N2-N3), 1.326(2) Å (N3-N4), and 1.342(2) Å (N4-N5) (reference bond lengths: C-N, 1.47 Å; N=N, 1.24 Å, and N-N, 1.47 Å).¹⁹ These values are in excellent agreement with those reported in the literature and are also quite uniform within the series of azotetrazolates **2-5**. Thus the structural parameters of the remaining three complexes will not be discussed here in more detail. What makes the crystal structures of **2-5** special is the presence of extended hydrogen-bonded networks in the solid state as illustrates in Figures 3, 6, 9, and 12. Hydrogen-bonded networks are quite typical for nearly all previously reported salt-like azotetrazolates.¹⁹ Figure 12 displays the packing diagram of [Co(NH₃)₅N₃](C₂N₁₀)·H₂O (**5**) as a typical example. It can be seen that the azotetrazolate anions form stacks interconnecting the [Co(NH₃)₅N₃]²⁺ cations. Six of the nitrogen atoms of the tetrazolyl rings are engaged in hydrogen bonding (see Table S12 in the ESI). Four of them (N7, N11, N12, N13, *cf.* Fig. 11) form N···H-N hydrogen bonds in the range of 3.013(2)-3.098(2) Å with NH₃ ligands in the [Co(NH₃)₅N₃]²⁺ cation, while the other two (N5, N8) form N···H-O hydrogen bonds (2.832(2) and 3.044(2) Å) with the water of crystallization. This water molecule obviously plays a central role in this crystal packing as it is also connected to two NH₃ ligands in the [Co(NH₃)₅N₃]²⁺ cation (O1···H-N18 and O1···H-14). Full details on the hydrogen bonding in all crystal structures reported here can be found in the ESI.

Nitrotetrazolates

Explosive materials comprising the 5-nitrotetrazolate anion, [CN₄NO₂]⁻ (= NT) play an important role in the chemistry and technology of nitrogen-rich energetic materials. Various salts²¹ and transition metal complexes²² containing nitrotetrazolate have already

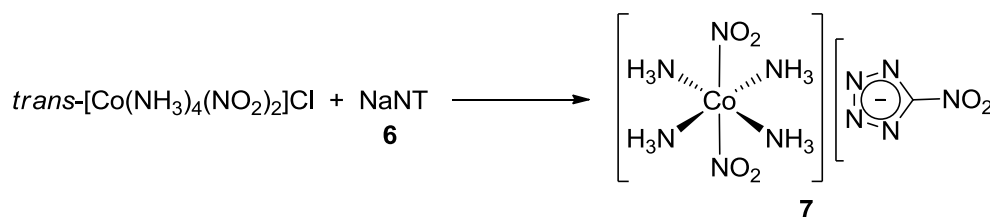
been prepared and tested. Recently, we described the synthesis, structural characterization, and energetic properties of $[\text{Me}_3\text{Sn}(\mu\text{-OH})\text{SnMe}_3(\mu\text{-OH})\text{SnMe}_3(\text{H}_2\text{O})][\text{NT}]$ as the first example of a metal-organic main-group metal nitrotetrazolate.²³ As mentioned above, perhaps the most promising candidate in this area is the recently reported copper(I) 5-nitrotetrazolate (= DBX-1).⁸ Manufacture of DBX-1 involves treatment of copper(II) chloride with sodium 5-nitrotetrazolate dihydrate (= NaNT, **6**) in the presence of sodium ascorbate as reducing agent. The synthesis of the important precursor NaNT (**6**) has been investigated in great detail.^{21e,24} For use in this study, NaNT (**6**) was prepared by a modified literature method according to Scheme 5.^{21e,25} The synthesis involves a modified Sandmeyer reaction (diazotation) of 5-aminotetrazole in the presence of copper(II) sulfate. In the first step, a pale blue, very explosive (especially when dry) intermediate of the composition $\text{Cu}(\text{NT})_2 \cdot \text{HNT}$ is formed, which is kept in a wet state and subsequently treated with warm sodium hydroxide solution to afford **6** as the dihydrate. While the dihydrate of NaNT is fairly insensitive and can be handled with the appropriate care, the anhydrous material is dangerously explosive.^{21e,25} Most recently, a significantly improved preparation of high-purity NaNT has been published, which avoids the handling of the potentially dangerous $\text{Cu}(\text{NT})_2 \cdot \text{HNT}$ intermediate.^{24c}



Scheme 5

In the course of the present study, the first cationic Werner-type cobalt(III) ammine complex containing 5-nitrotetrazolate as counterion, *i.e.* not directly coordinated to the central Co^{3+} ion (*cf.* Scheme 2), has been prepared according to Scheme 6. Combination of aqueous solutions of equimolar amounts of *trans*- $[\text{Co}(\text{NH}_3)_4(\text{NO}_2)_2]\text{Cl}$ and sodium 5-nitrotetrazolate dihydrate (= NaNT, **6**) afforded the novel cobalt(III) 5-nitrotetrazolate

derivative *trans*-[Co(NH₃)₄(NO₂)₂](NT)·H₂O (**7**) as orange, rectangular prismatic crystals in 64% yield according to Scheme 6.



Scheme 6

The IR spectrum of **7** is dominated by strong or medium bands attributable to the asymmetric (1550 cm⁻¹ (s)) and symmetric (1322 cm⁻¹ (s)) stretching vibrations and the deformation vibration (820 cm⁻¹ (m)) of the nitro group.^{24a} A ¹³C NMR spectrum, measured with a saturated solution in D₂O, showed the resonance of the single tetrazole ring carbon atom at δ 165.2 ppm, which differs only slightly from the values around δ 169 ppm reported for alkali metal 5-nitrotetrazolates.^{24a} The formation of *trans*-[Co(NH₃)₄(NO₂)₂](NT)·H₂O was also verified by a single-crystal X-ray diffraction study (Table 1, Figures 13-15). Compound **7** crystallizes in the tetragonal space group I₄-1 as the monohydrate.

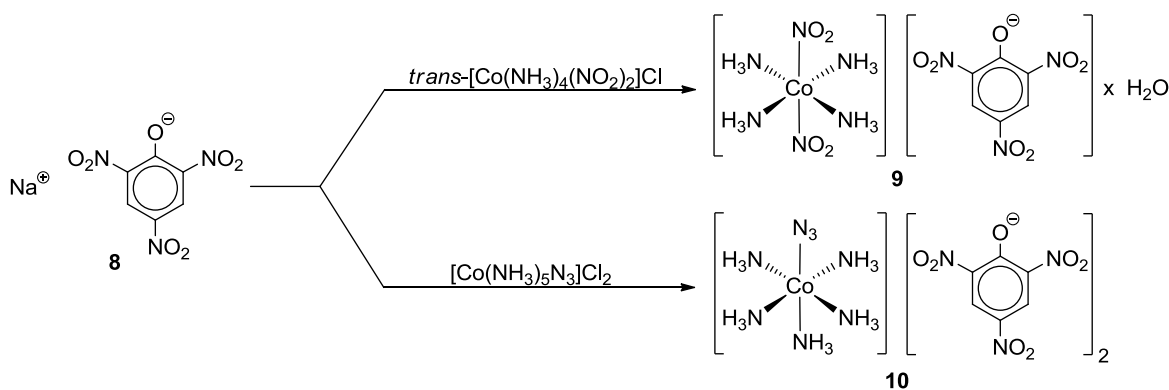
Figures 13-15 here

Similar to the azotetrazolates **2-5**, the individual ions in **7** show no unexpected structural features. With 1.331(3)-1.333(3) Å all five bonds within the tetrazole ring are identical and comparable to those in other salts comprising the 5-nitrotetrazolate anion.²⁴ The nitro group is highly symmetrical with N-O distances of 1.228(2) Å. As shown in Fig. 15, the crystal structure of **7** also comprises an extensive network of hydrogen bonds. The crystal structure consists of mutually linked *trans*-[Co(NH₃)₄(NO₂)₂]⁺ cations and 5-nitrotetrazolate anions. Details on the hydrogen-bridging can be found in Table S15 in the ESI. Notably, there are four intramolecular O···H-N hydrogen bonds in the range of 2.901(2)-3.380(2) Å *within* the *trans*-[Co(NH₃)₄(NO₂)₂]⁺ cation between coordinated NH₃ molecules and oxygen atoms of the nitro groups. The water of crystallization connects cation and anion *via* one N···H-O hydrogen bond to N5 in the tetrazole ring (N5···H-O3 2.960(2) Å) and one O···H-N hydrogen bond to one of the coordinated NH₃ molecules

(O3...H-N3 3.082(2) Å). Finally, there is one direct N...H-N hydrogen bond between N6 in the tetrazole ring and one of the NH₃ ligands (N6...H-N3B 3.156(2) Å).

Picrates

It is well established that anhydrous picric acid tends to be unstable, and its impact and friction sensitivities are higher than those of TNT. Nevertheless, numerous organic and inorganic picrate salts have been prepared and their crystal structures studied.²⁶ Shreeve et al. have reported mono and bridged azolium picrates as energetic salts.^{7c,27} Picrates of Werner-type cationic cobalt(III) ammine complexes have also been frequently reported, but mainly with respect to studying their solubility and/or crystallinity. A typical example is a publication by Ephraim in which the solubility of various complex cobalt(III) picrates was studied.²⁸ These included *e.g.* [Co(NH₃)₅H₂O](picrate)₃, [Co(NH₃)₄(H₂O)₂](picrate)₃, [Co(NH₃)₅X](picrate)₂ (X = Cl, Br, I, NO₂, NO₃) as well as *cis*- and *trans*-[Co(NH₃)₄(NO₂)₂](picrate). Although the energetic properties of these compounds were not explicitly studied, it was noted that combustion analyses proved difficult due to regularly occurring explosions.²⁸ In the course of our study, two Werner-type complex cobalt(III) picrates were prepared as outlined in Scheme 7.



Scheme 7

Sodium picrate, prepared *in situ* by neutralizing picric acid with NaOH, was subsequently treated with either *trans*-[Co(NH₃)₄(NO₂)₂]Cl or [Co(NH₃)₅N₃]Cl₂. The products, *trans*-[Co(NH₃)₄(NO₂)₂](picrate)·H₂O (9) and [Co(NH₃)₅N₃](picrate)₂ (10), which precipitated from the concentrated aqueous solutions, were isolated by filtration and purified by recrystallization from a minimum amount of hot water. Complex 9 was

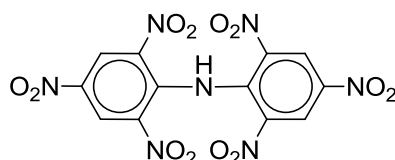
obtained in 57% yield as large, orange blocks (Fig. 16), whereas complex **10** (87% yield) formed dark red rods (Fig. 19). Complex **9** has already been mentioned in the early work by Ephraïm.²⁸ Notably, this author described no less than four different crystal forms for this compound, depending on the recrystallization conditions: 1. yellow, hair-like crystals of several centimeters length; 2. glistening needles and platelets of rhombic shape; and 3. thick, heavy prism-like crystals. As can be seen in Fig. 16, we apparently obtained only the latter sort of crystals. Both the IR and ¹³C NMR data of **9** and **10** showed the typical values for the picrate anion. Both complexes were also structurally characterized by single-crystal X-ray diffraction (Table 1, Figures 16-21 and ESI).

Figures 16-21 here

The dinitro complex **9** crystallizes in the triclinic space group P-1. Figure 17 shows that the unit cell contains two cations, two anions and two water molecules. All bond lengths and angles in the individual components show no significant deviations from those reported in the literature.^{21,22} Figure 18 illustrates the supramolecular hydrogen-bonded crystal structure of **9**. The two water molecules, the ammine and nitro ligands of the *trans*-[Co(NH₃)₄(NO₂)₂]⁺ cation, and the nitro groups of the picrate anion all participate in the complex hydrogen-bonded network. The azido derivative **10** also crystallizes in the triclinic space group P-1, but has no water of crystallization. As can be seen in Figure 21, all ligands in the [Co(NH₃)₅N₃]²⁺ dication as well as the oxygen atoms of the phenoxide and nitro groups of the picrate anions participate in the N-H...O and N-H...N hydrogen bonds forming the supramolecular network (see ESI for full details).

Dipicrylamides

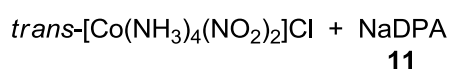
Dipicrylamine (2,2',4,4',6,6'-hexanitrodiphenylamine, Scheme 8) combines several interesting structural features in that it contains six nitro groups which are flexible and can interact and adjust in the crystal lattice and in that it has a secondary amine group which can be deprotonated with alkali and alkaline earth metal hydroxides to form water soluble salts. In the resulting dipicrylamide anion (= DPA⁻), partial delocalization of the negative charge mediated by the aromatic rings is possible, which may facilitate coordination of the oxygen atoms of the nitro groups with suitable metal ions.²⁹



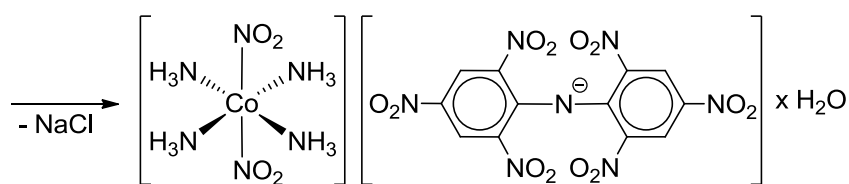
Scheme 8

The ammonium salt of dipicrylamine, also known as Aurantia or Imperial Yellow, was discovered in 1874 by Gnehm and used as a yellow colorant for leather, wool, and silk until the early 20th century.³⁰ However, this use has been terminated due to the highly toxic and explosive nature of dipicrylamine.³¹ Dipicrylamine can also be used for the extraction of K^+ ions from sea bittern (a mixture of K^+ , Na^+ , and Mg^{2+}).^{32a} A related study carried out with a mixture of K^+ , Rb^+ , and Cs^+ revealed that the Cs^+ ion shows maximum selectivity towards DPA^- .^{32b} In fact, it has been reported that DPA^- can be used for the recovery of Cs^+ from radioactive wastes.³³ Only recently, the structural chemistry of alkali metal and alkaline earth metal as well as ammonium and azolium dipicrylamides has been investigated in detail. All these compounds display interesting hydrogen-bonded supramolecular structures in the solid state.²⁹

A Werner-type cobalt(III)-ammine complex of dipicrylamide was prepared according to Scheme 9 by treatment of *trans*- $[Co(NH_3)_4(NO_2)_2]Cl$ with 1 equiv. of *in situ*-prepared NaDPA (Scheme 9).



11



12

Scheme 9

Compound **12** was isolated in 53% yield in the form of orange, somewhat irregular blocks (Fig. 22). Both the IR spectrum and the elemental analysis indicated the formation

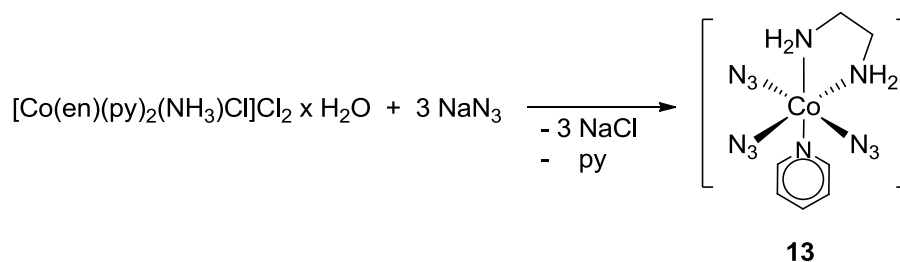
of a hydrated species. An X-ray structural analysis confirmed the presence of the monohydrate *trans*-[Co(NH₃)₄(NO₂)₂](DPA)·H₂O (**12**) (Table 1, Figures 22-24).

Figures 22-24 here

Complex **12** crystallizes in the triclinic space group P-1. Figure 23 shows that the unit cell contains two crystallographically independent *trans*-[Co(NH₃)₄(NO₂)₂]⁺ cations and two DPA anions. The structural parameters (bond lengths and angles) in the dipicrylamide anions show only little differences when compared to those reported for the alkali, alkaline earth metals, ammonium or azolium salts of dipicrylamide.²⁹

The non-electrolyte complex *mer*-[Co(en)(py)(N₃)₃] (**13**)

Well known from the early studies on Werner-type cobalt(III) azido complexes is the non-electrolyte complex *mer*-[Co(NH₃)₃(N₃)₃]. The dark green, highly explosive material was first reported by Linhard and Weigel in 1950.¹⁸ This dark blue-green compound was found to form on various occasions when aquo- or azido-cobalt(III) ammine complexes are treated with excess sodium azide. The best synthetic method (94% yield) involves reaction of [Co(NH₃)₄(H₂O)₂](ClO₄)₃ with a large excess of NaN₃. Almost quantitative yields of *mer*-[Co(NH₃)₃(N₃)₃] can also be obtained in a simple manner by air-oxidation of a mixture of CoSO₄·7H₂O, (NH₄)₂SO₄ and NaN₃ in aqueous ammonia solution. It was noted that even small amounts of *mer*-[Co(NH₃)₃(N₃)₃] could lead to dangerous detonations upon heating grinding or impact. Detonations were even observed when larger crystals were ground under water.¹⁸ We now succeeded in the preparation and structural characterization of a new non-electrolyte complex formally derived from cobalt(III) triazide. To an aqueous solution of [Co(en)(py)₂(NH₃)Cl]Cl₂·H₂O³⁴ was added a *ca.* 20-fold excess of NaN₃ and the reaction mixture was heated on a steam bath until all the water had evaporated. During the reaction a color change from red to dark green occurred and the smell of pyridine became evident. Unreacted sodium azide was easily removed by repeated washing of the dark residue with cold water. After drying in air, the new non-electrolyte complex *mer*-[Co(en)(py)(N₃)₃] (**13**) was isolated in 93% yield as a dark blue-green crystalline solid (Scheme 10). As expected for a non-electrolyte complex, the solubility of **13** in water is very low, although the resulting very dilute solutions show an intense dark green color.



Scheme 10

This material was found to be unpredictably and dangerously explosive. It should thus be prepared and handled only in very small amounts and with utmost care. Compound **13** was even too dangerously explosive to be fully characterized by IR and elemental analysis. Attempts to obtain an IR spectrum of **13** were unsuccessful, because an explosion occurred upon grinding of ca. 2-3 mg with ca. 100 mg of KBr. As a derivative of the unknown cobalt triazide, compound **13** has a high nitrogen content of ca. 52% which certainly accounts for its highly explosive nature. Despite all the dangerous properties, it proved possible to structurally verify the nature of **13** by single-crystal X-ray diffraction. Even this task was met by major obstacles. Recrystallization from hot water was impossible due to the low solubility of **13**. Thus small portions of the originally obtained materials were spread on glass sample holders and searched under a microscope for suitable crystals. It turned out that most of the crystals were strongly intergrown, with most of them having the appearance of Christmas trees. Notably, like *mer*-[Co(NH₃)₃(N₃)₃],¹⁸ the crystals of **13** were blue-green dichroitic. Finally, a small blue-green crystal fragment was found to be suitable for X-ray diffraction (Table 1 and Figs. 25 and 26; see also ESI). As is shown in Fig. 25, the three azide ligands in **13** are arranged in meridional positions as in the parent non-electrolyte *mer*-[Co(NH₃)₃(N₃)₃].¹⁸ The Co-N bond lengths to the two opposing azide ligands are identical (Co(1)-N(4) 1.946(2), Co(1)-N(2) 1.948(2) Å), whereas that to the azide ligands *trans* to an amino group of ethylenediamine is slightly longer (Co(1)-N(7) 1.962(2) Å). The packing diagram of **13** shows a polymeric hydrogen-bonded network (Fig. 26). In this case, supramolecular association of the complex molecules occurs only *via* N⋯H-N hydrogen bonds between terminal azide nitrogen atoms and the NH₂ groups of the coordinated ethylenediamine.

Energetic properties of the new cobalt(III) complexes

The methods for testing primary explosives have just been summarized in a very informative short review article by Mehta et al.³⁵ In addition to a simple combustion test, the impact and friction sensitivity were tested according to established BAM methods using a BAM drophammer and BAM friction tester (Figures 27 and 28).^{1c} Initial test results of the energetic properties of seven cobalt(III) complexes prepared in the course of this study are listed in Table 2. Not included in the Table is the dipicrylamide derivative **12** because it was found to be insensitive to both impact and friction. Also not included is the non-electrolyte complex *mer*-[Co(en)(py)(N₃)₃] (**13**). This complex was found to be too sensitive to be handled safely. All tests were repeated six times over a storage period of 30 days. These test series revealed no significant deviations from the originally measured values, indicating a good long-term stability of the title compounds. Moreover, the drophammer tests were repeated in the temperature range of -55 °C to 20 °C in order to get an impression of the low-temperature performance. For example, hunting and sports ammunition is expected to show good performance at temperatures down to -25 °C, whereas military ammunition should function even at temperatures as low as -55 °C. Thus, drophammer tests of all seven compounds were performed at -55°C, -30 °C, 0 °C, 10 °C, and 20 °C. In all cases these tests revealed an unexpected increase of the impact sensitivity by 1-2 J at temperatures below 0 °C which remained constant to -55 °C. Although this effect may be small, it is significant and reproducible. In summary, all energetic cobalt(III) complexes showed a stable performance over the entire temperature range of 75 degrees.

As shown in Table 2, all four azotetrazolates were found to be impact-sensitive. For the azido complex **5**, the drophammer test revealed that this compound is very sensitive to impact. The value of 3.5-5 J for **5** is comparable to those found for Pb(N₃)₂ (3.0-6.5 J) or lead(II) styphnate (2.5-5.0 J),^{1b} so that **5** can be classified as a primary explosive. This is not surprising in view of the fact that **5** has the highest nitrogen content (68.47%) of all compounds studied. In contrast, all four compounds showed only low sensitivity to friction, with compound **3** being more or less insensitive. Three of the samples (**3**, **4**, and **5**) exploded upon ignition with a flame, while a sample of compound **2** showed only deflagration.

The values for the 5-nitrotetrazolate complex **7** show that with a low value of 5 J the impact sensitivity of this compound is comparable to that of lead(II) azide. Thus this compound can also be classified as a primary explosive. This, combined with a good oxygen balance Ω of -13.67% and a relatively high nitrogen content of ~44%, makes

compound **7** a promising candidate for further testing. In the flame test, a sample of complex **7** showed deflagration.

Both picrates **9** and **10** can also be classified as primary explosives. With values of 8 (**9**) and 7.5 (**10**) they are slightly less sensitive to impact than lead(II) styphnate or lead(II) azide and at the same time much less friction-sensitive. However, both samples exploded upon ignition with a flame. With 27.10% (**9**) and 30.53% (**10**) the nitrogen content is still quite good, while the negative oxygen balance can be explained by the high carbon content of these picrates. In view of the easy accessibility of these and related complexes, such Werner-type cobalt(III) picrates certainly merit further investigation as alternative energetic materials. As mentioned above, the dipicrylamide complex **12** was found to be insensitive to both impact and friction (>40 J and >360 N, resp.). This is in agreement with recent findings by Zhou et al. who reported that the impact sensitivities of various ammonium and azolium dipicrylamides are in the range of that of the secondary explosive TNT.^{33c} In view of the known high toxicity of dipicrylamine,^{30,33c} this energetic anion does not appear to be a useful alternative for lead(II)-containing explosives anyway. In contrast, the non-electrolyte complex **13** could not be safely tested as even minor samples of 2-3 mg could explode in an unpredictable manner.

CONCLUSIONS

In summarizing the work reported here, we have investigated the synthesis, structures, and energetic properties of a series of novel Werner-type cobalt(III) ammine complexes comprising the energetic anions azotetrazolate, nitrotetrazolate, picrate and dipicrylamide in the outer sphere, *i.e.* not directly coordinated to the central Co^{3+} ion. The title compounds are accessible in a straightforward manner by simple salt metathesis reactions using readily available starting materials. Also easily prepared, though unpredictably and dangerously explosive, is the new dark blue-green non-electrolyte complex *mer*-[Co(en)(py)(N₃)₃] (**13**). The crystal structures of all new complexes are characterized by extensive hydrogen-bonded networks. In addition to the structural authentication by X-ray diffraction, the energetic properties of these explosive Werner-type cobalt(III) complexes have been tested. The tests revealed that several of the title complexes can be classified as primary explosives. Particularly promising in that respect are the compounds **5** and **10** which contain the [Co(NH₃)₅N₃]⁺ cation, as well as the nitrotetrazolate *trans*-[Co(NH₃)₄(NO₂)₂](NT)·H₂O (**7**). All tested samples showed excellent long-term storage and low-temperature stability.

EXPERIMENTAL SECTION

General Procedures

The starting materials $[\text{Co}(\text{NH}_3)_5\text{NO}_2]\text{Cl}_2$,³⁶ $[\text{Co}(\text{NH}_3)_4(\text{py})\text{NO}_2]\text{Cl}_2$ (py = pyridine),³⁷ *trans*- $[\text{Co}(\text{NH}_3)_4(\text{NO}_2)_2]\text{Cl}$,³⁸ $[\text{Co}(\text{NH}_3)_5\text{N}_3]\text{Cl}_2$,¹³ $[\text{Co}(\text{en})(\text{py})_2(\text{NH}_3)\text{Cl}]\text{Cl}_2 \cdot \text{H}_2\text{O}$,³⁴ $\text{Na}_2\text{C}_2\text{N}_{10} \cdot 5\text{H}_2\text{O}$, (**1**),²⁰ and NaN_3 (**6**)^{21e,25} were prepared according to published procedures. Picric acid, dipicrylamide and NaN_3 were obtained from commercial sources and used as received. IR (KBr) spectra were measured using a Perkin-Elmer FT-IR 2000 spectrometer. NMR spectra were recorded on a DPX 400 (Bruker) or AVANCE 600 (Bruker-Biospin) spectrometer. Microanalyses of the compounds were performed using a Leco CHNS 023 apparatus.

Single crystal X-ray crystallography

The intensity data of **2**, **3**, **4**, **5**, **7**, **10**, and **13** were collected on a Stoe IPDS 2T diffractometer with MoK_α radiation. The data were collected with the Stoe XAREA³⁹ program using ω -scans. The space groups were determined with the XRED32³⁹ program. The intensity data of **9** and **12** were registered on an Oxford Diffraction Nova A diffractometer using mirror-focused CuK_α radiation. Absorption corrections were applied using the multi-scan method. The structures were solved by direct methods (SHELXS-97) and refined by full matrix least-squares methods on F^2 using SHELXL-97.⁴⁰

Cautionary note

*All compounds described in this study are sensitive energetic materials which should only be synthesized and manipulated on a small (<250 mg) scale using proper safety equipment including thick leather gloves and jackets, face shields and blast screens, ear plugs and plastic or Teflon laboratory equipment. Care should be taken not to extensively dry or heat these materials as the anhydrous complexes are even more sensitive and prone to explosion on manipulation. In particular, the non-electrolyte complex **13** is extremely sensitive to impact and friction and should be handled with utmost care.*

Synthesis of pentaammine-nitro-cobalt(III)-azotetrazolate-dihydrate, $[\text{Co}(\text{NH}_3)_5\text{NO}_2](\text{C}_2\text{N}_{10}) \cdot 2\text{H}_2\text{O}$ (2**).** To a solution of pentaammin-nitro-cobalt(III)-chloride (0.20 g, 0.77 mmol) in a minimum amount of warm water (40 °C, *ca.* 10 mL) were added

0.24 g (0.80 mmol) $\text{Na}_2\text{C}_2\text{N}_{10}\cdot 5\text{H}_2\text{O}$ (**1**) in 10 mL water. Undisturbed standing at r.t. for 24 h afforded orange, clear yellow crystals suitable for X-ray diffraction (Fig. 1). Yield: 0.23 g (74%). M.p. > 290 °C (violent dec.). $^{13}\text{C}\{^1\text{H}\}$ NMR (100.6 MHz, $\text{DMSO-}d_6$, 20 °C): δ 173.1 ($[\text{N}_4\text{C-N=N-CN}_4]^{2-}$). IR (KBr, cm^{-1}): ν_{max} 3431 (vs), 2106 (m), 1628 (m), 1452 (s), 1402 (s), 1202 (m), 1164 (m), 1065 (m), 1042 (m), 866 (m), 775 (m), 740 (s), 564 (s). Anal. Calcd for $\text{C}_2\text{H}_{19}\text{CoN}_{16}\text{O}_4$ (392.23 g/mol): C, 6.16; H, 4.91; N, 57.43. Found: C, 6.02; H, 4.76; N, 56.78.

Synthesis of *trans*-tetraammine-nitro-pyridine-cobalt(III)-azotetrazolate-dihydrate, $[\text{Co}(\text{NH}_3)_5\text{NO}_2](\text{C}_2\text{N}_{10})\cdot 2\text{H}_2\text{O}$ (3**).** In a similar manner, *trans*-tetraammine-nitro-pyridine-cobalt(III)-dichloride (0.11 g, 0.34 mmol) was allowed to react with 0.10 g (0.33 mmol) $\text{Na}_2\text{C}_2\text{N}_{10}\cdot 5\text{H}_2\text{O}$ (**1**) in ca. 15 mL water. Undisturbed standing at r.t. for 24 h afforded yellow-orange crystals suitable for X-ray diffraction (Fig. 4). Yield: 0.21 g (46%). M.p. > 260-263 °C (violent dec.). $^{13}\text{C}\{^1\text{H}\}$ NMR (100.6 MHz, $\text{DMSO-}d_6$, 20 °C): δ 173.1 ($[\text{N}_4\text{C-N=N-CN}_4]^{2-}$), 158.2 (2 C, py), 135.1 (1 C, py), 127.2 (2 C, py). IR (KBr, cm^{-1}): ν_{max} 3544 (vs), 3430 (vs), 2445 (m), 2108 (m), 1626 (s), 1506 (m), 1452 (m), 1403 (s), 1203 (m), 1165 (m), 1096 (m), 1064 (m), 1043 (m), 776 (m), 741 (s), 566 (s). Anal. Calcd for $\text{C}_7\text{H}_{21}\text{CoN}_{16}\text{O}_4$ (452.28 g/mol): C, 18.59; H, 4.68; N, 49.55. Found: C, 20.01; H, 4.20; N, 49.37.

Synthesis of bis[*trans*-tetraammine-dinitro-cobalt(III)]-azotetrazolate, *trans*- $[\text{Co}(\text{NH}_3)_4(\text{NO}_2)_2]_2(\text{C}_2\text{N}_{10})$ (4**).** Reaction between *trans*- $[\text{Co}(\text{NH}_3)_4(\text{NO}_2)_2]\text{Cl}$ (0.20 g, 0.79 mmol) and **1** (0.11 g, 0.37 mmol) afforded 0.19 g (87%) **4** as clear yellow single-crystals suitable for X-ray diffraction (Fig. 7). M.p. > 250 °C (violent dec.). $^{13}\text{C}\{^1\text{H}\}$ NMR (100.6 MHz, $\text{DMSO-}d_6$, 20 °C): δ 173.1 ($[\text{N}_4\text{C-N=N-CN}_4]^{2-}$). IR (KBr, cm^{-1}): ν_{max} 3331 (m), 2445 (m), 2106 (m), 1628 (m), 1452 (m), 1375 (s), 1202 (m), 1164 (m), 1065 (m), 1042 (m), 866 (m), 775 (m), 740 (m), 692 (m), 564 (m), 464 (m). Anal. Calcd for $\text{C}_2\text{H}_{24}\text{Co}_2\text{N}_{22}\text{O}_8$ (602.22 g/mol): C, 3.99; H, 4.02; N, 51.17. Found: C, 3.68; H, 4.10; N, 53.01.

Synthesis of pentaammine-azido-cobalt(III)-azotetrazolate-monohydrate, $[\text{Co}(\text{NH}_3)_5\text{N}_3](\text{C}_2\text{N}_{10})\cdot \text{H}_2\text{O}$ (5**).** Reaction between $[\text{Co}(\text{NH}_3)_5\text{N}_3]\text{Cl}_2$ (0.35 g, 1.37 mmol) and **1** (0.41 g, 1.37 mmol) afforded a deep purple reaction mixture from which **5** crystallized after 1d in the form of purple, rod-like single-crystals which were suitable for X-ray diffraction (Fig. 10). Yield: 0.44 g (88%). M.p. 256 °C (violent dec.). $^{13}\text{C}\{^1\text{H}\}$ NMR

(100.6 MHz, DMSO-*d*₆, 20 °C): δ 173.2 ([N₄C-N=N-CN₄]²⁻). IR (KBr, cm⁻¹): ν_{\max} 3275 (vs), 2346 (s), 2050 (vs), 1631 (vs), 1558 (vs), 1385 (vs), 1315 (vs), 1152 (s), 1041 (s), 830 (vs), 771 (vs), 734 (vs), 557 (s). Anal. Calcd for C₂H₁₇CoN₁₈O (368.21 g/mol): C, 6.52; H, 4.65; N, 68.47. Found: C, 6.66; H, 4.02; N, 70.99.

Synthesis of *trans*-tetraammine-dinitro-cobalt(III)-5-nitrotetrazolate-monohydrate, *trans*-[Co(NH₃)₄(NO₂)₂](NT)·H₂O (7). Similar treatment of *trans*-[Co(NH₃)₄(NO₂)₂]Cl (0.15 g, 0.58 mmol) with **6** (0.10 g, 0.58 mmol) in a minimum amount of water (*ca.* 10 mL) afforded 0.10 g (64%) of **7** as orange, block-like single crystals with a somewhat milky appearance (Fig. 13), which were nevertheless suitable for X-ray diffraction. M.p. *ca.* 230 °C (violent dec.). ¹³C{¹H} NMR (100.6 MHz, D₂O, 20 °C): δ 165.2 (1 C, NO₂-CN₄). IR (KBr, cm⁻¹): ν_{\max} 3306 (s), 1622 (m), 1550 (s), 1385 (vs), 1322 (vs), 1282 (s), 819 (m), 623 (m), 503 (m). Anal. Calcd for CH₁₄CoN₁₁O₇ (351.12 g/mol): C, 3.42; H, 4.02; N, 43.88. Found: C, 3.09; H, 3.88; N, 42.90.

Synthesis of *trans*-tetraammine-dinitro-cobalt(III)-picrate-monohydrate, *trans*-[Co(NH₃)₄(NO₂)₂](picrate)·H₂O (9). Reaction of *trans*-[Co(NH₃)₄(NO₂)₂]Cl (0.10 g, 0.39 mmol) with sodium picrate (0.39 mmol, made *in situ* from 0.09 g of picric acid and NaOH) in a minimum amount of water (*ca.* 10 mL) first gave an orange precipitate. Recrystallization of the crude product from hot water (*ca.* 8 mL) afforded 0.12 g (66%) of **9** as orange, block-like single-crystals (Fig. 16) which were suitable for X-ray diffraction. M.p. *ca.* 285 °C (violent dec.). ¹³C{¹H} NMR (20°C, D₂O, 100 MHz): δ (100.6 MHz, D₂O, 20 °C): δ 160.9 (1 C; C1), 141.9 (2 C; C2, C6), 125.3 (2 C; C3, C5), 124.3 (1 C; C4). IR (KBr, cm⁻¹): ν_{\max} 3629 (m), 3522 (s), 3308 (vs), 1635 (vs), 1556 (vs), 1510 (s), 1474 (s), 1435 (s), 1400 (vs), 1366 (vs), 1321 (vs), 1250 (s), 1162 (s), 1086 (s), 946 (m), 929 (m), 911 (m), 818 (s), 789 (m), 747 (m), 709 (s), 602 (m), 507 (m). Anal. Calcd for C₆H₁₆CoN₉O₁₂ (465.18 g/mol): C, 15.49; H, 3.47; N, 27.10. Found: C, 15.45; H, 3.67; N, 27.10.

Synthesis of pentaammine-azido-cobalt(III)-dipicrate, [Co(NH₃)₅N₃](picrate)₂ (10). Analogous reaction of [Co(NH₃)₅N₃]Cl₂ (0.10 g, 0.39 mmol) with sodium picrate (0.78 mmol, made *in situ* from 0.18 g of picric acid and NaOH) in a minimum amount of water (*ca.* 10 mL) afforded a dark red-purple precipitate. Recrystallization from hot water (*ca.* 10 mL) gave 0.22 g (87%) of **10** as clear, dark red, rod-like single-crystals (Fig. 19)

which were suitable for X-ray diffraction. M.p. *ca.* 195 °C (violent dec.). $^{13}\text{C}\{^1\text{H}\}$ NMR (100.6 MHz, D_2O , 20 °C): δ 160.9 (1 C; C1), 141.8 (2 C; C2, C6), 125.3 (2 C; C3, C5), 124.0 (1 C; C4). IR (KBr, cm^{-1}): ν_{max} 3271 (vs), 2056 (vs), 1639 (s), 1554 (s), 1433 (m), 1368 (m), 1320 (vs), 1285 (s), 1163 (m), 1083 (m), 910 (m), 843 (s), 747 (m), 712 (m), 662 (m), 579 (m), 463 (m). Anal. Calcd for $\text{C}_{12}\text{H}_{19}\text{CoN}_{14}\text{O}_{14}$ (642.30 g/mol): C, 22.44; H, 2.98; N, 30.53. Found: C, 23.40; H, 3.04; N, 30.48.

Synthesis of *trans*-tetraammine-dinitro-cobalt(III)-dipicrylamide-mono-hydrate, *trans*- $[\text{Co}(\text{NH}_3)_4(\text{NO}_2)_2](\text{DPA})\cdot\text{H}_2\text{O}$ (12**).** To a suspension of dipicrylamine (1.0 g, 2.27 mmol) in 10 mL water were added two pellets of solid NaOH to give a dark red solution of NaDPA.³³ A concentrated solution of *trans*- $[\text{Co}(\text{NH}_3)_4(\text{NO}_2)_2]\text{Cl}$ (0.58 g, 2.27 mmol) in 10 mL water was added and the dark red reaction mixture was allowed to stand undisturbed for 1 week. This produced bright orange crystals of **12** in 53% yield (0.42 g). M.p. > 350 °C (dec.). IR (KBr, cm^{-1}): ν_{max} 3436 (m), 3360 (m), 2430 (w), 1790 (m), 1665 (m), 1587 (s), 1498 (m), 1383 (vs), 1320 (vs), 1300 (s), 1262 (m), 1982 (m), 910 (w), 836 (m), 768 (w), 740 (w), 719 (m), 509 (w). Anal. Calcd for $\text{C}_{24}\text{H}_{38}\text{Co}_2\text{N}_{26}\text{O}_{35}$ (1368.59 g/mol): C, 21.06; H, 2.80; N, 26.61. Found: C, 21.00; H, 2.55; N, 24.98.

Synthesis of *mer*-triazido-ethylenediamine-pyridine-cobalt(III), *mer*- $[\text{Co}(\text{en})(\text{py})(\text{N}_3)_3]$ (13**).** 0.30 g (0.72 mmol) of $[\text{Co}(\text{en})(\text{py})_2(\text{NH}_3)\text{Cl}]\text{Cl}_2\cdot\text{H}_2\text{O}$ were dissolved in 25 mL H_2O and 1.00 g (15.38 mmol) NaN_3 were added as solid. The reaction mixture was heated in a steam bath without stirring until the water had evaporated completely. During this time the color of the solution changed from red to dark green and the smell of pyridine became evident. The residue was swirled five times with 20 mL- portions of water and the aqueous solutions were carefully decanted in order to redissolve and separate the unreacted NaN_3 . After drying the residue in air, 0.22 g (93%) of **13** were obtained as dark blue-green, feather-like crystals. Due to its highly explosive nature, neither melting/decomposition point, IR, nor elemental analysis data could be obtained for **13**. Anal. Calcd for $\text{C}_7\text{H}_{13}\text{CoN}_{12}$ (324.19 g/mol): C, 25.93; H, 4.04; N, 51.85.

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Table 1. Crystallographic Data for **2, 3, 4, 5, 7, 9, 10, 12, and 13.**

	2 (ip103)	3 (ip123)	4 (ip99)	5 (bl154a)	7 (ip115)	9 (magd08)	10 (ip152)	12 (magd20)	13 (fred01)
Empirical formula	C ₂ H ₁₉ CoN ₁₆ O ₄	C ₇ H ₂₁ CoN ₁₆ O ₄	C ₂ H ₂₄ Co ₂ N ₂₂ O ₈	C ₂ H ₁₇ CoN ₁₈ O	CH ₁₄ CoN ₁₁ O ₇	C ₆ H ₁₆ CoN ₉ O ₁₂	C ₁₂ H ₁₉ CoN ₁₄ O ₁₄	C ₂₄ H ₃₈ Co ₂ N ₂₆ O ₃₅	C ₇ H ₁₃ CoN ₁₂
<i>a</i> (Å)	12.510(3)	7.3460(15)	7.1228(14)	12.067(2)	7.8926(11)	6.8481(3)	7.0622(14)	12.6374(3)	6.2607(12)
<i>b</i> (Å)	7.4092(15)	9.907(2)	7.3402(15)	7.4493(15)	7.8926(11)	13.4859(5)	12.888(3)	13.7368(3)	6.8540(12)
<i>c</i> (Å)	17.702(4)	13.308(3)	10.138(2)	15.764(3)	19.125(4)	18.5624(6)	13.299(3)	15.6996(4)	15.006(2)
α (°)	90	71.60(3)	96.49(3)	90	90	75.421(3)°	98.33(3)	113.942(2)	84.600(14)
β (°)	100.07(3)	83.33(3)	105.79(3)	90.94(3)	90	89.746(3)°	104.80(3)	90.678(2)	79.945(15)
γ (°)	90	78.54(3)	95.10(3)	90	90	85.199(3)°	90.52(3)	89.527(2)	85.933(15)
<i>V</i> (Å ³)	1615.5(6)	899.2(3)	502.79(17)	1416.8(5)	1191.3(3)	1653.03(11)	1156.6(4)	2490.71(10)	630.22(19)
<i>Z</i>	4	2	1	4	4	4	2	2	2
formula weight	390.21	452.28	602.22	368.21	351.12	465.18	642.30	1368.59	324.19
space group	P2 ₁ /c	P -1	P-1	P2 ₁ /n	I4 ₁ md	P-1	P-1	P-1	P-1
<i>T</i> (°C)	-140	-140	-140	-140	-140	-173	-140	-173	-173
λ (Å)	0.71073	0.71073	0.71073	0.71073	0.71073	1.54184	0.71073	1.54184	1.54184
<i>D</i> _{calcd} (g cm ⁻³)	1.605	1.671	1.989	1.726	1.958	1.869	1.844	1.825	1.709
μ (mm ⁻¹)	1.111	1.011	1.740	1.250	1.502	8.991	0.845	6.434	10.826
<i>F</i> (000)	808	468	308	760	720	952	656	1396	332
data/restraints/parameters	3973 / 35 / 243	4440 / 4 / 305	2457 / 0 / 202	2764 / 0 / 267	855 / 2 / 67	6845 / 0 / 521	6204 / 33 / 374	10298 / 0 / 825	2563 / 6 / 197
goodness-of-fit on <i>F</i> ²	1.096	1.193	1.099	0.974	1.208	1.051	1.040	1.039	0.955
<i>R</i> (<i>F</i> _o or <i>F</i> _o ²)	0.0577	0.0611	0.0325	0.0226	0.0232	0.0566	0.0512	0.0374	0.0386
<i>R</i> _w (<i>F</i> _o or <i>F</i> _o ²)	0.1341	0.1282	0.0876	0.0573	0.0566	0.1553	0.0998	0.0956	0.0879

Table 2. Energetic properties of **2**, **3**, **4**, **5**, **7**, **9**, and **10**.

Compound	2	3	4	5	7	9	10
M (g/mol)	392.23	452.28	602.22	368.21	351.12	465.18	642.30
ρ , g/cm ³	1.605	1.671	1.989	1.726	1.958	1.869	1.844
Impact sensitivity/J	8	> 10	8	3.5-5	5	8	7.5
Friction sensitivity (N)	> 216	> 360	> 216	216	> 216	> 216	> 216
Combustion test	deflagr.	explosion	explosion	explosion	deflagr.	explosion	explosion
N (%)	57.43	49.55	51.17	68.47	43.88	27.10	30.53
Ω (%)	- 55.87	- 82.64	- 26.56	- 49.97	- 13.67	-30.95	-26.15

Figures

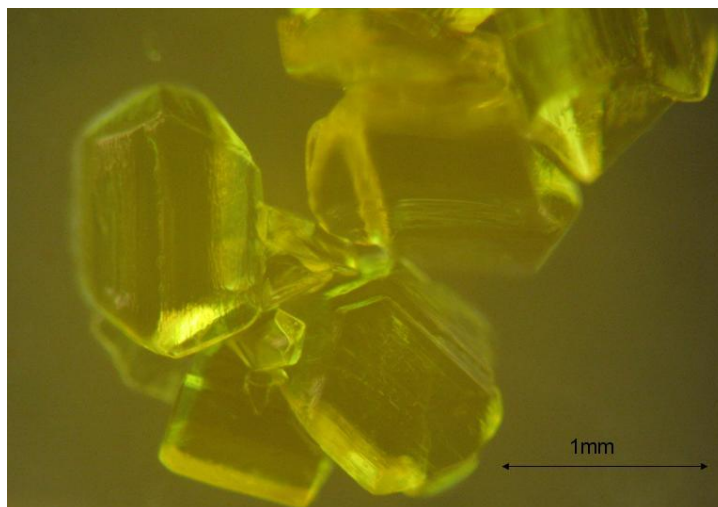


Fig. 1 Single-crystals of $[\text{Co}(\text{NH}_3)_5\text{NO}_2](\text{C}_2\text{N}_{10})\cdot 2\text{H}_2\text{O}$ (**2**).

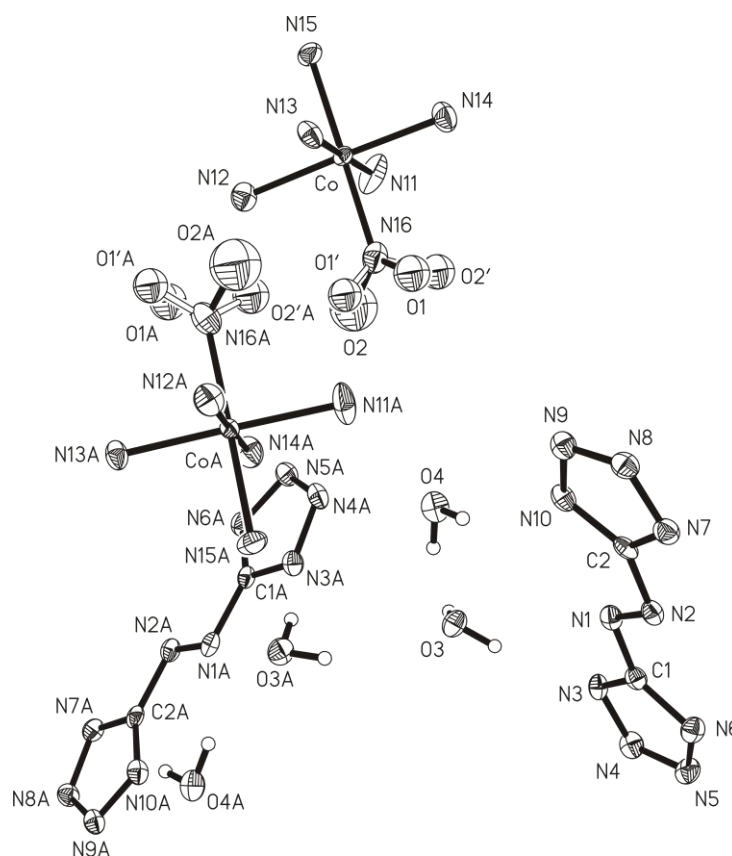


Fig. 2 ORTEP drawing of $[\text{Co}(\text{NH}_3)_5\text{NO}_2](\text{C}_2\text{N}_{10})\cdot 2\text{H}_2\text{O}$ (**2**) at the 50% probability level. Selected bond lengths (\AA) and angles ($^\circ$): Co-N(13) 1.941(3), Co-N(16) 1.946(4), Co-N(14) 1.949(3), Co-N(11) 1.956(3), Co-N(12) 1.965(3), Co-N(15) 1.969(3), N(1)-N(2) 1.252(4), N(1)-C(1) 1.410(5), N(2)-C(2) 1.407(5), N(2)-N(1)-C(1) 112.6(3), N(1)-N(2)-C(2) 111.7(3).

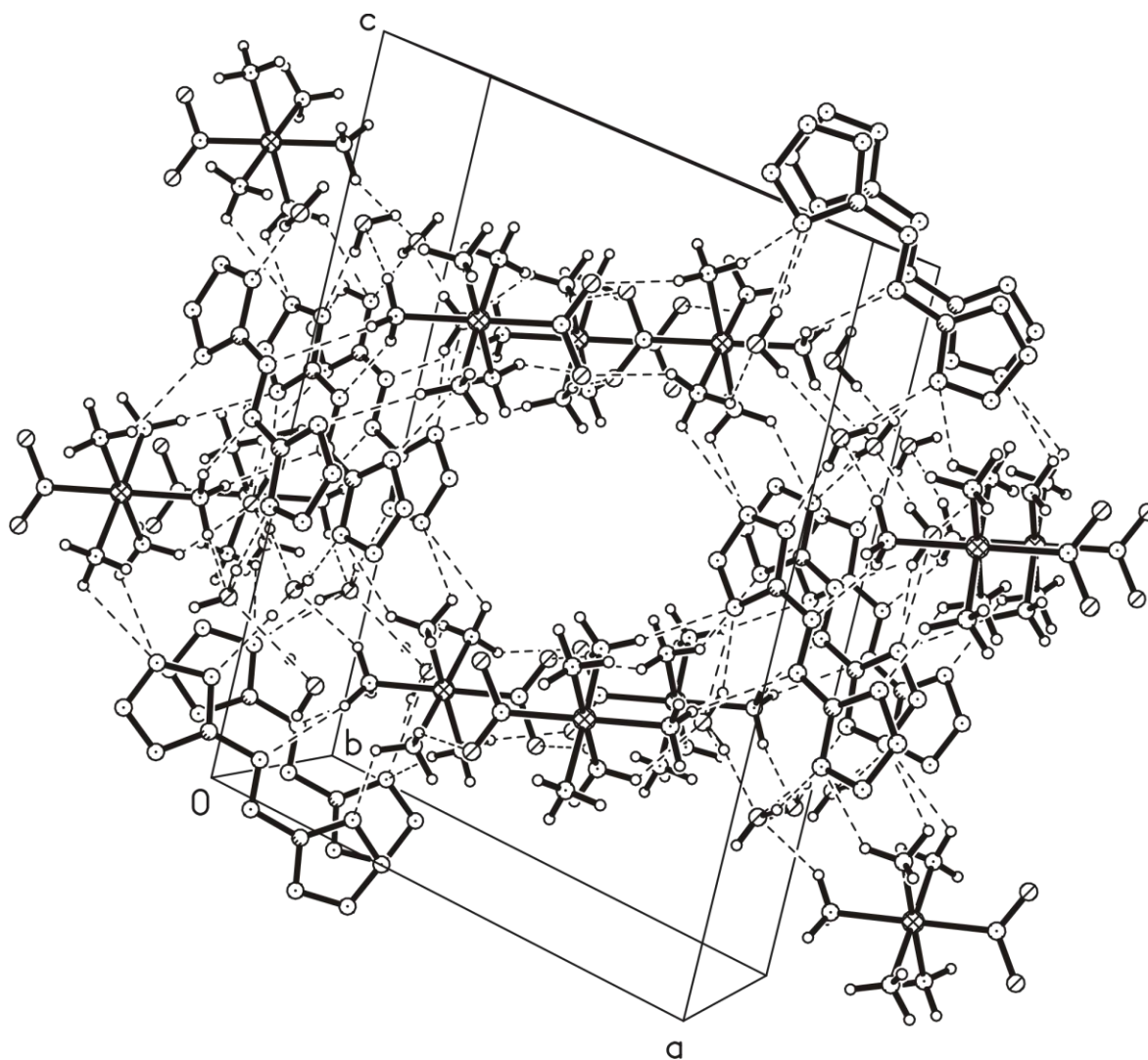


Fig. 3 Packing diagram of $[\text{Co}(\text{NH}_3)_5\text{NO}_2](\text{C}_2\text{N}_{10}) \cdot 2\text{H}_2\text{O}$ (**2**). View along the *b* axis.

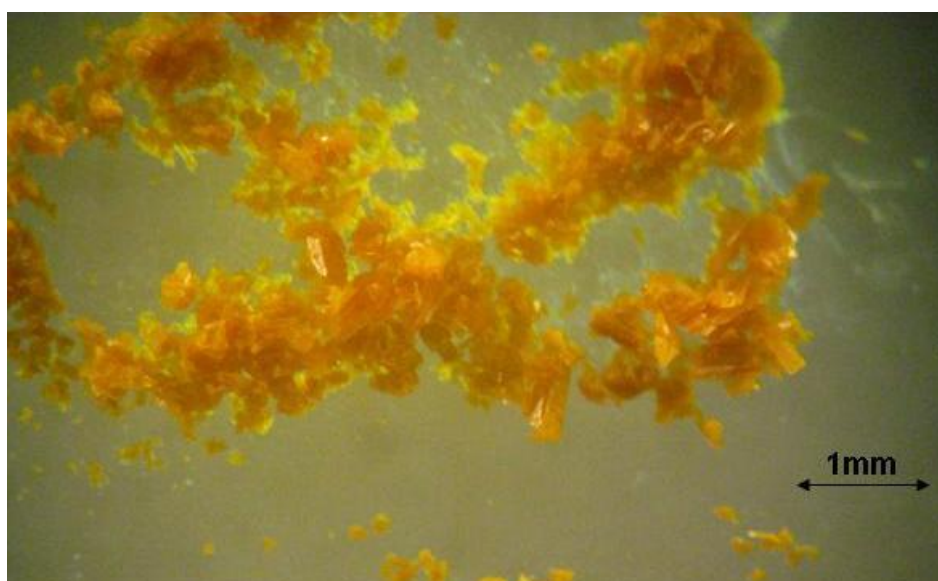


Fig. 4 Single-crystals of $[\text{Co}(\text{NH}_3)_4(\text{py})\text{NO}_2](\text{C}_2\text{N}_{10})\cdot 2\text{H}_2\text{O}$ (**3**).

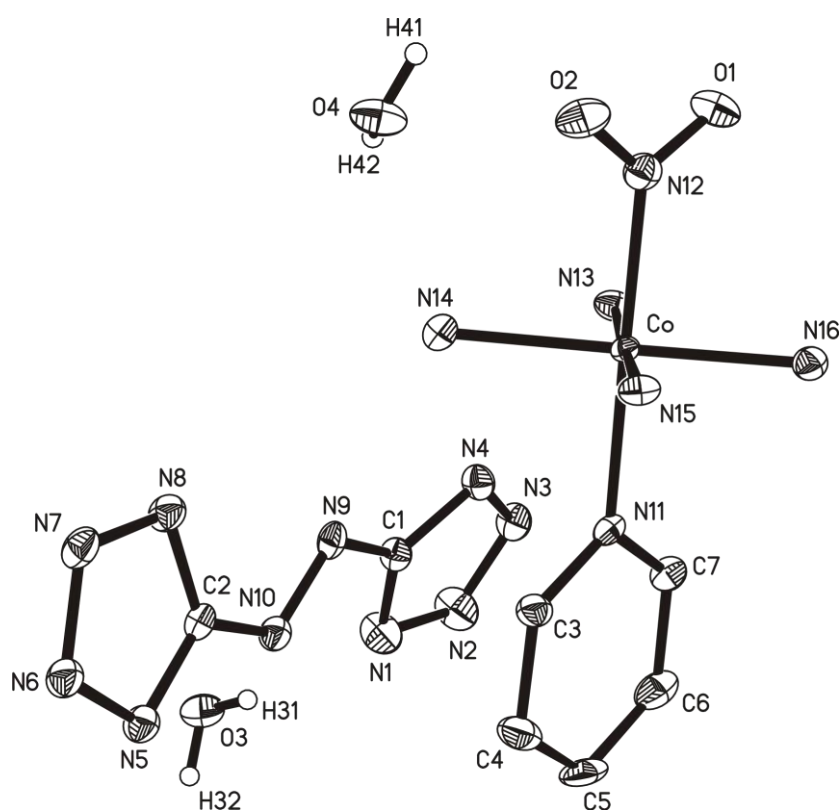


Fig. 5 ORTEP drawing of $[\text{Co}(\text{NH}_3)_4(\text{py})\text{NO}_2](\text{C}_2\text{N}_{10})\cdot 2\text{H}_2\text{O}$ (**3**) at the 50% probability level. Selected bond lengths (Å) and angles (°): Co-N(11) 1.996(3), Co-N(12) 1.899(3), Co-N(13) 1.961(3), Co-N(14) 1.961(3), Co-N(15) 1.961(3), Co-N(16) 1.956(3), N(9)-N(10) 1.261(5), N(9)-C(1) 1.408(5), N(10)-C(2) 1.399(5), N(12)-Co-N(11) 178.98(15), N(10)-N(9)-C(1) 112.8(3), N(9)-N(10)-C(2) 113.4(3).

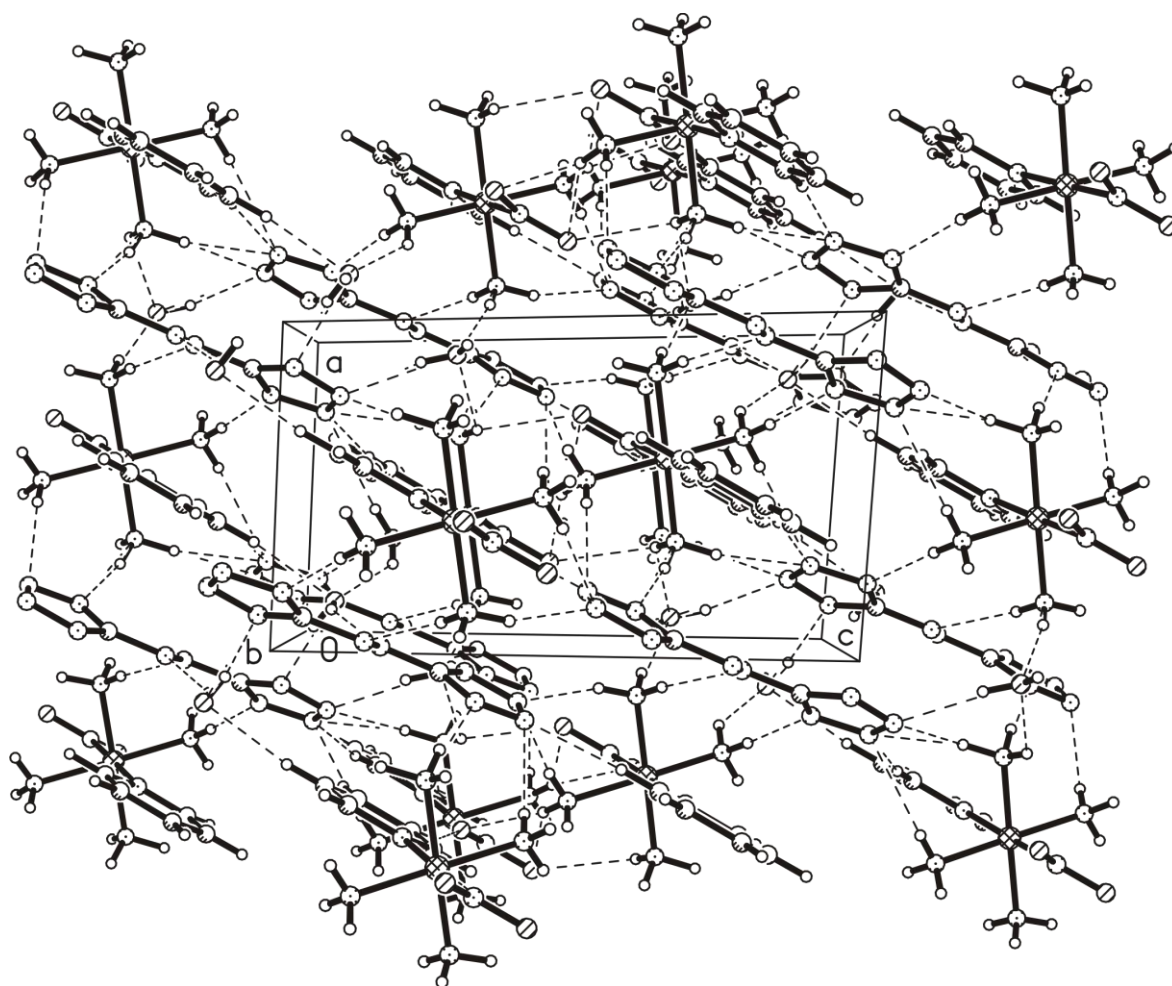


Fig. 6 Packing diagram of $[\text{Co}(\text{NH}_3)_4(\text{py})\text{NO}_2](\text{C}_2\text{N}_{10}) \cdot 2\text{H}_2\text{O}$ (**3**). View along the *b* axis.

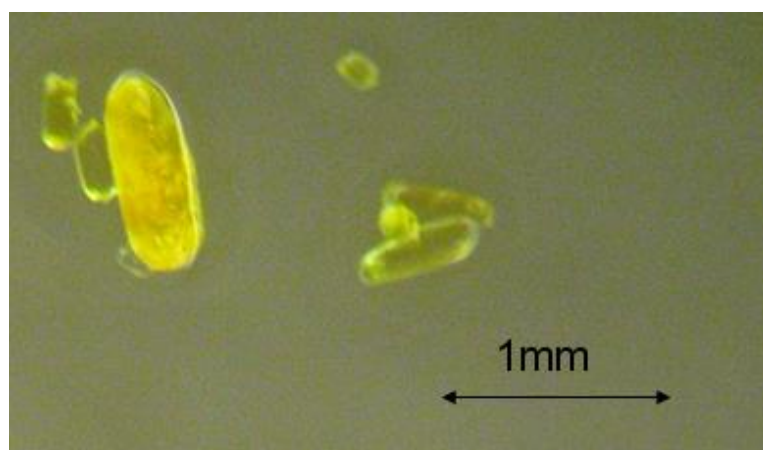


Fig. 7 Single-crystals of *trans*-[Co(NH₃)₄(NO₂)₂]₂(C₂N₁₀) (**4**).

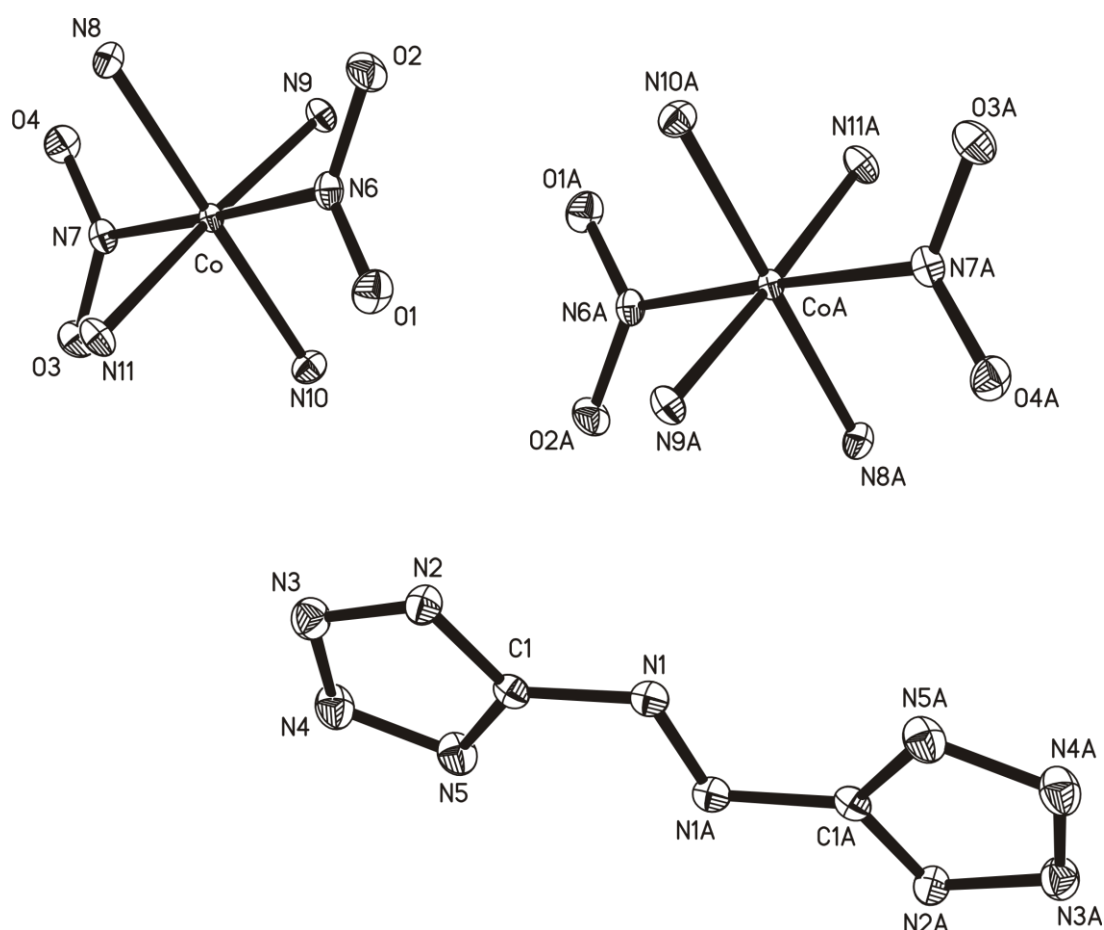


Fig. 8 ORTEP drawing of *trans*-[Co(NH₃)₄(NO₂)₂]₂(C₂N₁₀) (**4**) at the 50% probability level. Selected bond lengths [Å] and angles [°]: Co-N(7) 1.9319(15), Co-N(6) 1.9360(15), Co-N(9) 1.9453(16), Co-N(8) 1.9514(17), Co-N(11) 1.9532(16), Co-N(10) 1.9555(18), C(1)-N(1) 1.405(2), N(1)-N(1)#1 1.251(3), N(1)#1-N(1)-C(1) 113.37(18). (#1 -x+1,-y+1,-z+2)

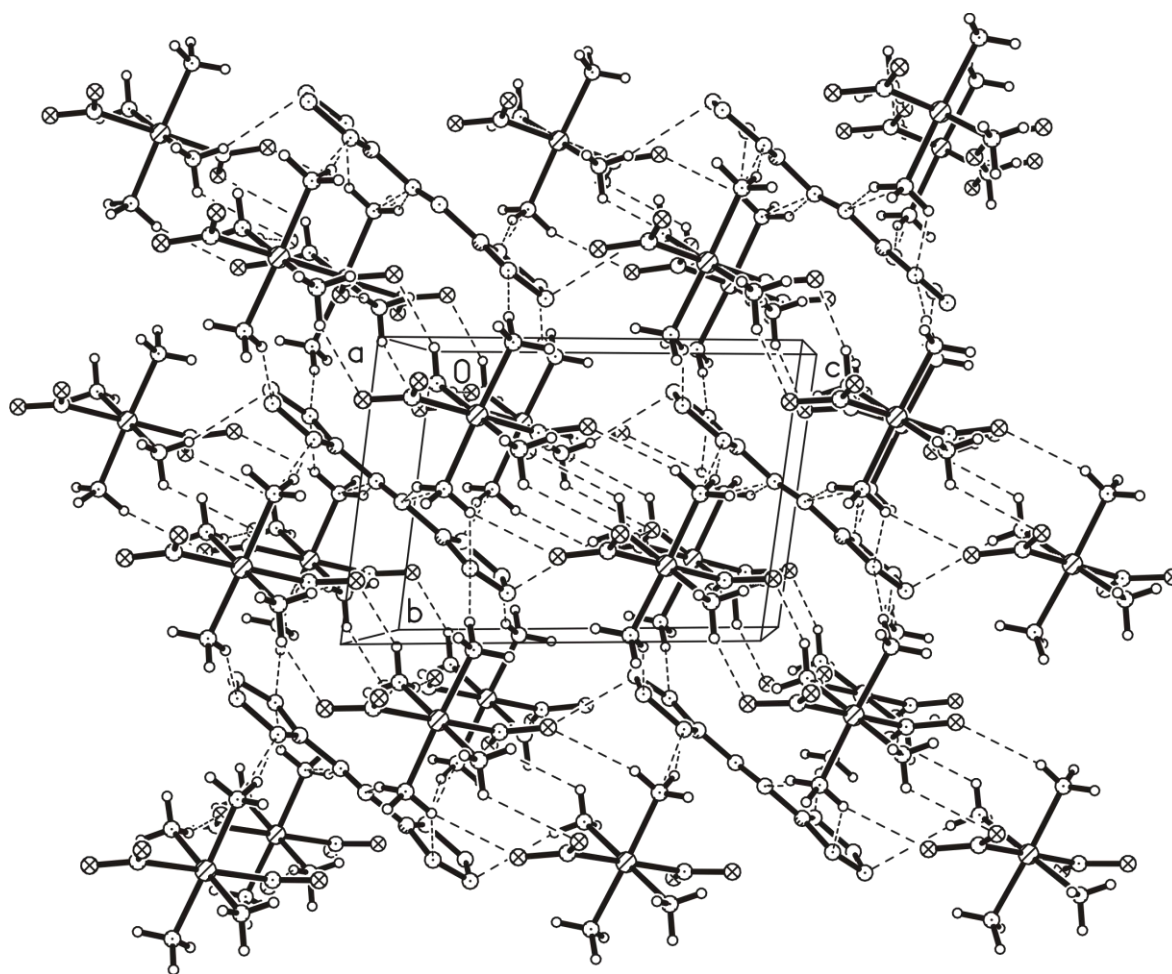


Fig. 9 Packing diagram of $trans-[Co(NH_3)_4(NO_2)_2]_2(C_2N_{10})$ (**4**). View along the a axis.

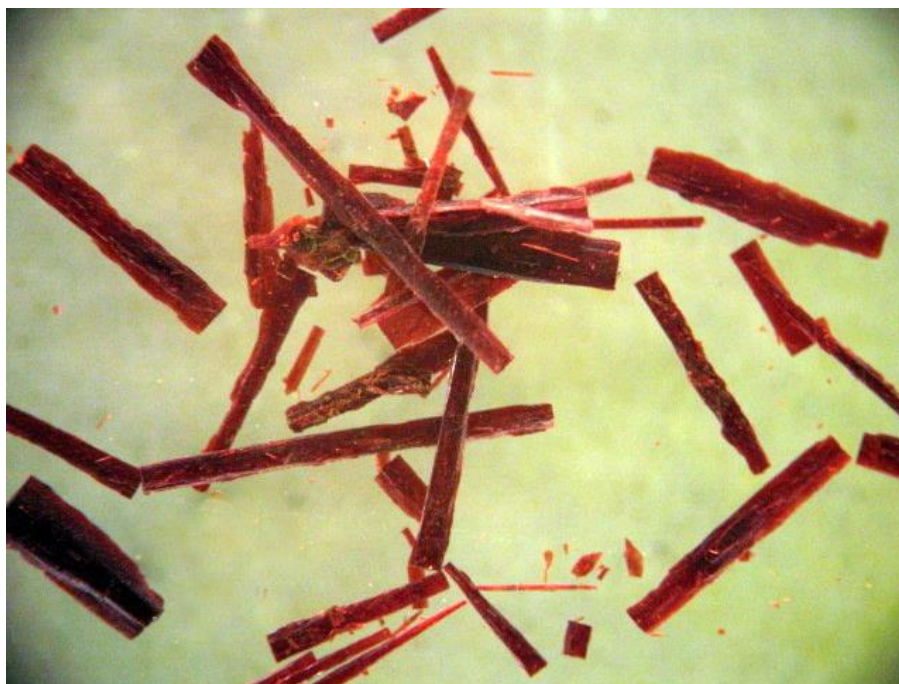


Fig. 10 Single-crystals of $[\text{Co}(\text{NH}_3)_5\text{N}_3](\text{C}_2\text{N}_{10})\cdot\text{H}_2\text{O}$ (**5**).

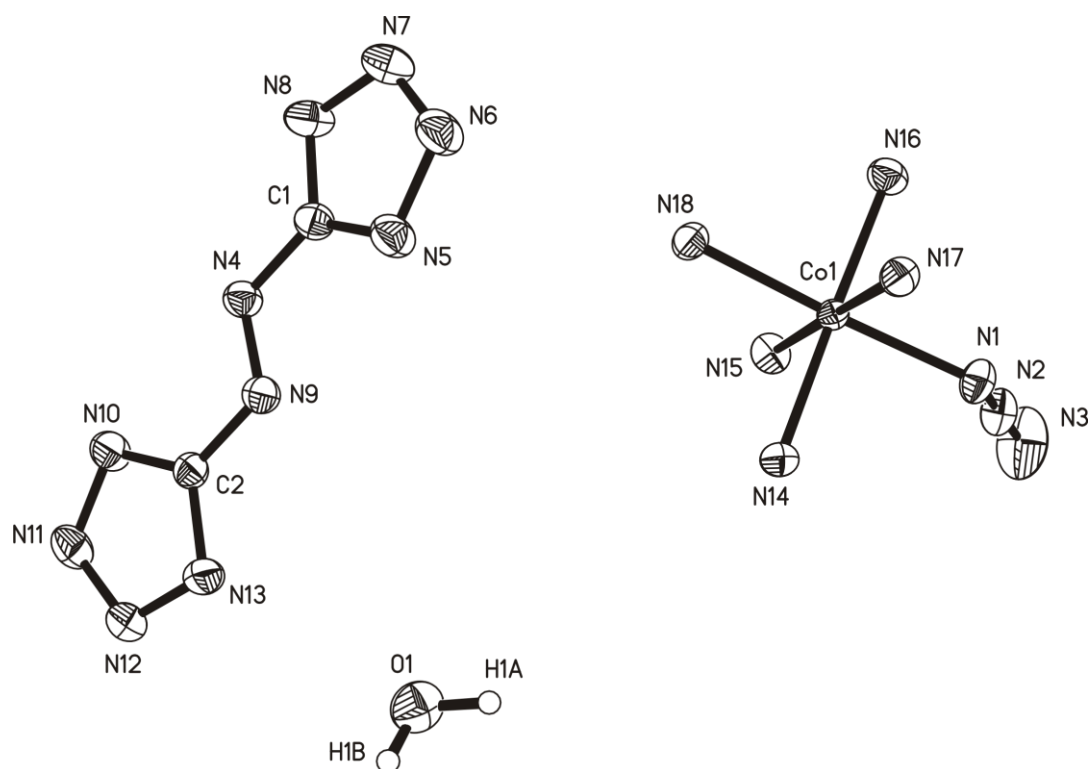


Fig. 11 ORTEP drawing of $[\text{Co}(\text{NH}_3)_5\text{N}_3](\text{C}_2\text{N}_{10})\cdot\text{H}_2\text{O}$ (**5**) at the 50% probability level. Selected bond lengths [\AA] and angles [$^\circ$]: Co(1)-N(14) 1.9496(13), Co(1)-N(1) 1.9550(14), Co(1)-N(17) 1.9566(13), Co(1)-N(18) 1.9637(14), Co(1)-N(16) 1.9758(14), Co(1)-N(15) 1.9804(14), N(4)-N(9) 1.2553(18), N(4)-N(9)-C(2) 113.09(12), N(9)-N(4)-C(1) 112.34(13).

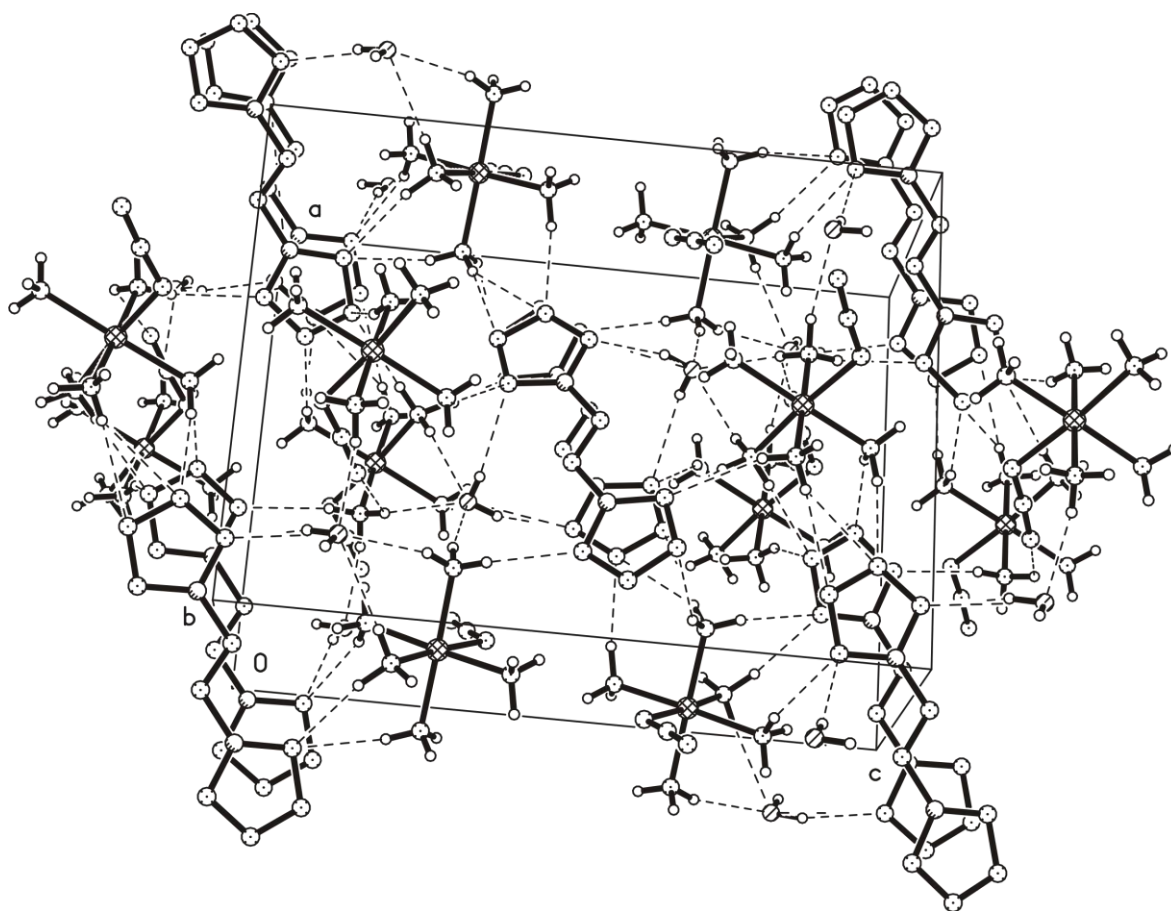


Fig. 12 Packing-diagram of $[\text{Co}(\text{NH}_3)_5\text{N}_3](\text{C}_2\text{N}_{10})\cdot\text{H}_2\text{O}$ (**5**). View along the *b* axis.

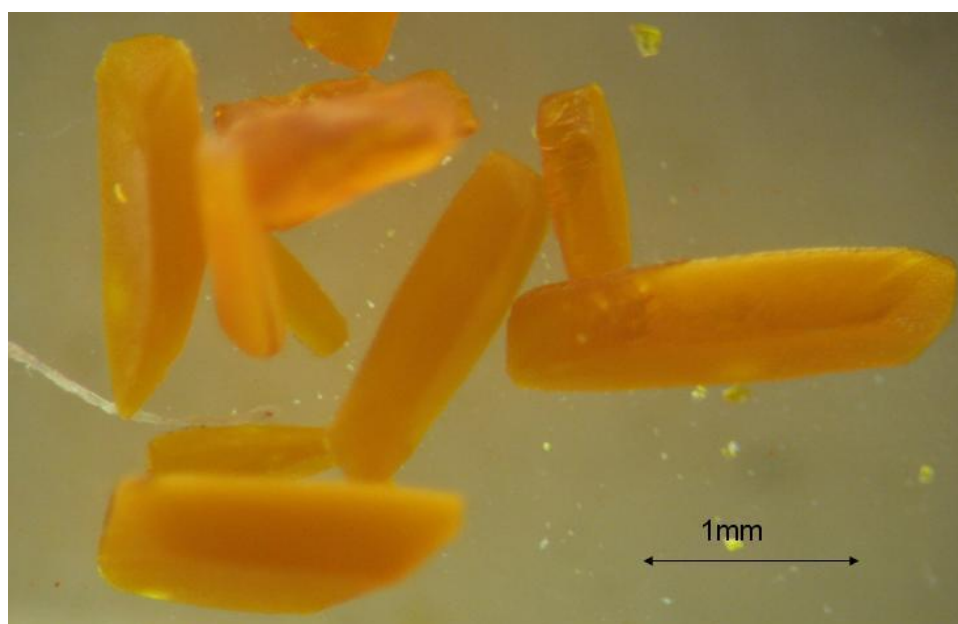


Fig. 13 Single-crystals of *trans*-[Co(NH₃)₄(NO₂)₂](NT)·H₂O (**7**).

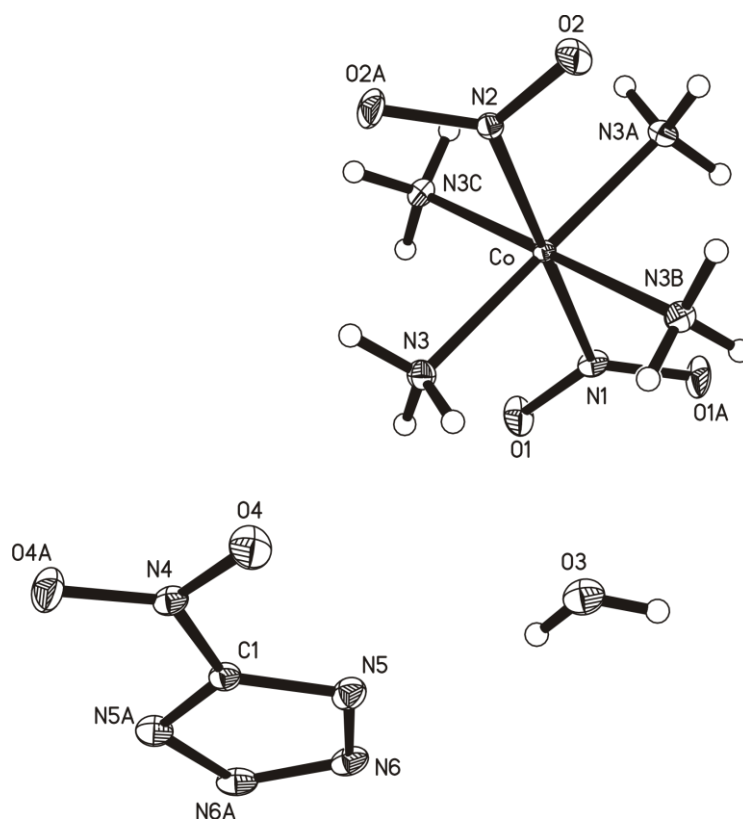


Fig. 14 ORTEP drawing of **7** *trans*-[Co(NH₃)₄(NO₂)₂](NT)·H₂O (**7**) at the 50% probability level. Selected bond lengths [Å] and angles [°]: Co-N(1) 1.937(3), Co-N(2) 1.942(4), Co-N(3) 1.9537(10), Co-N(3)#2 1.9537(10), Co-N(3)#3 1.9537(10), Co-N(3)#4 1.9537(10), C(1)-N(4) 1.438(6), O(4)-N(4)-C(1) 117.44(16). (#1 -x+1,-y+0,z+0 ; #2 -x+0,y+0,z+0 ; #3 x,-y+1,z ; #4 -x+0,-y+1,z+0)

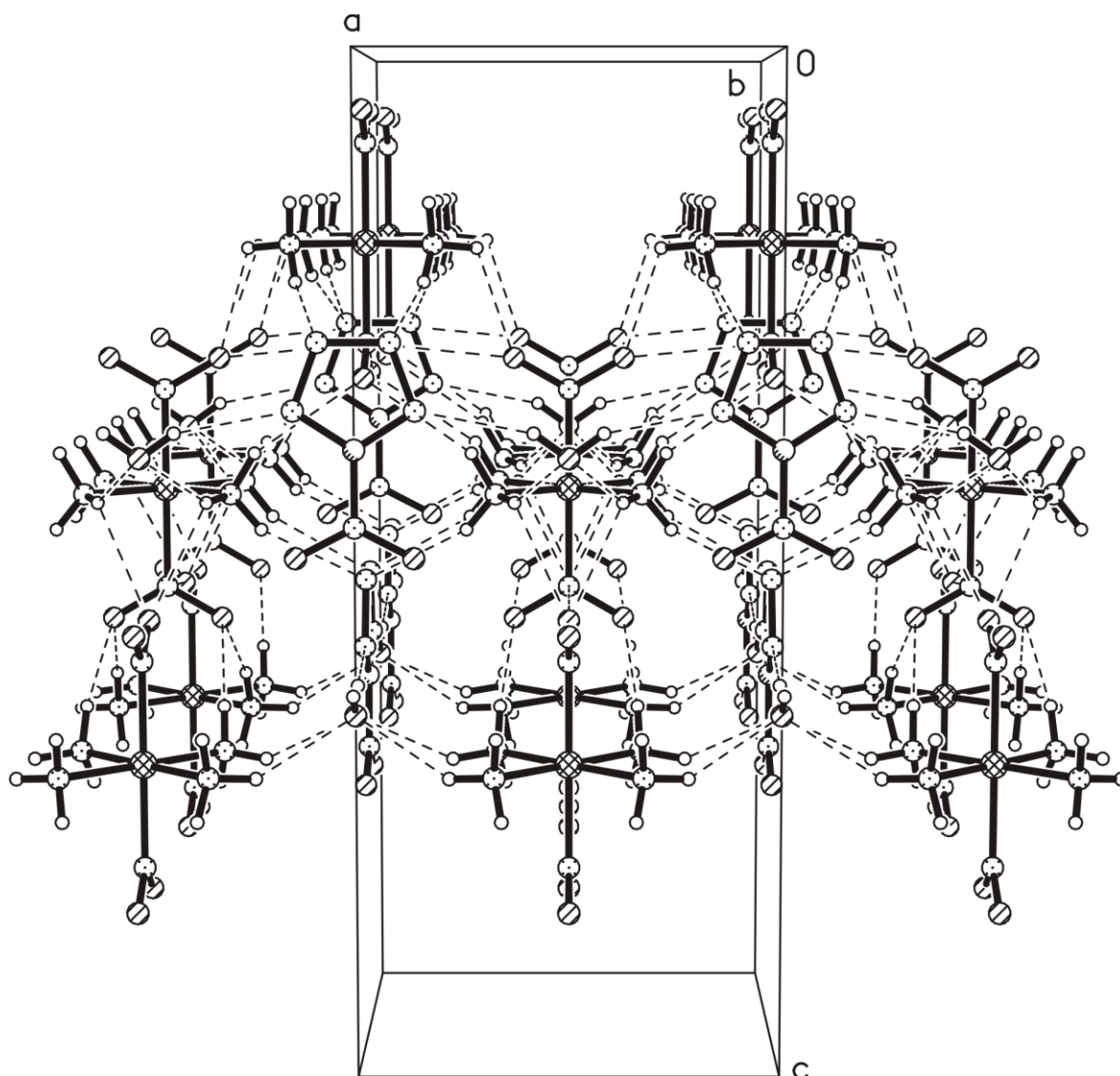


Fig. 15 Packing diagram of *trans*-[Co(NH₃)₄(NO₂)₂](NT)·H₂O (7). View along the *b* axis.

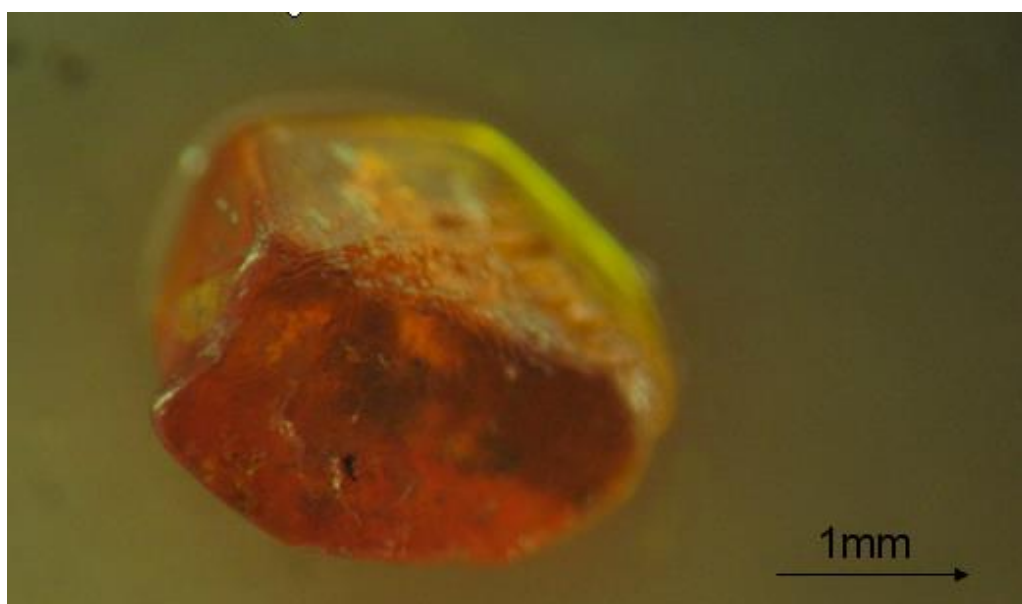


Fig. 16 Single-crystal of *trans*-[Co(NH₃)₄(NO₂)₂](picrate)·H₂O (**9**).

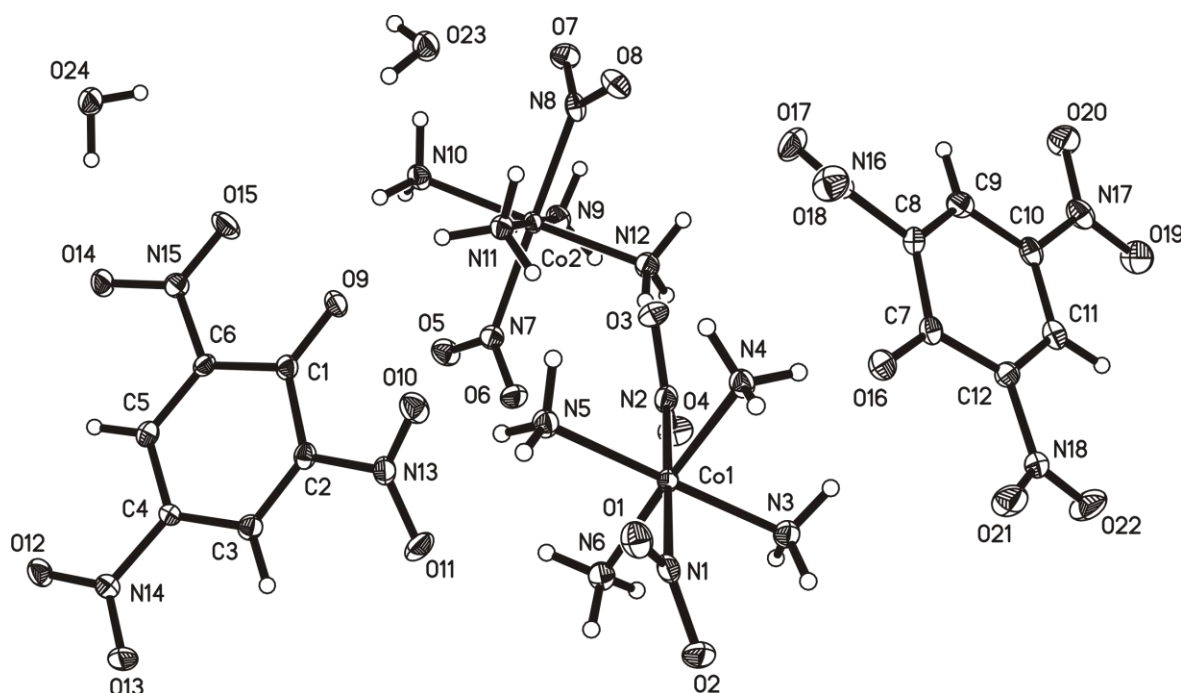


Fig. 17 ORTEP drawing of *trans*-[Co(NH₃)₄(NO₂)₂](picrate)·H₂O (**9**) at the 50% probability level. Selected bond lengths [Å] and angles [°]: Co(1)-N(2) 1.931(3), Co(1)-N(1) 1.943(3), Co(1)-N(3) 1.956(3), Co(1)-N(4) 1.962(3), Co(1)-N(6) 1.965(3), Co(1)-N(5) 1.967(3), Co(2)-N(8) 1.936(3), Co(2)-N(7) 1.942(3), Co(2)-N(12) 1.958(3), Co(2)-N(11) 1.959(3), Co(2)-N(10) 1.968(3), Co(2)-N(9) 1.973(3), N(2)-Co(1)-N(1) 179.39(13), N(8)-Co(2)-N(7) 179.05(13).

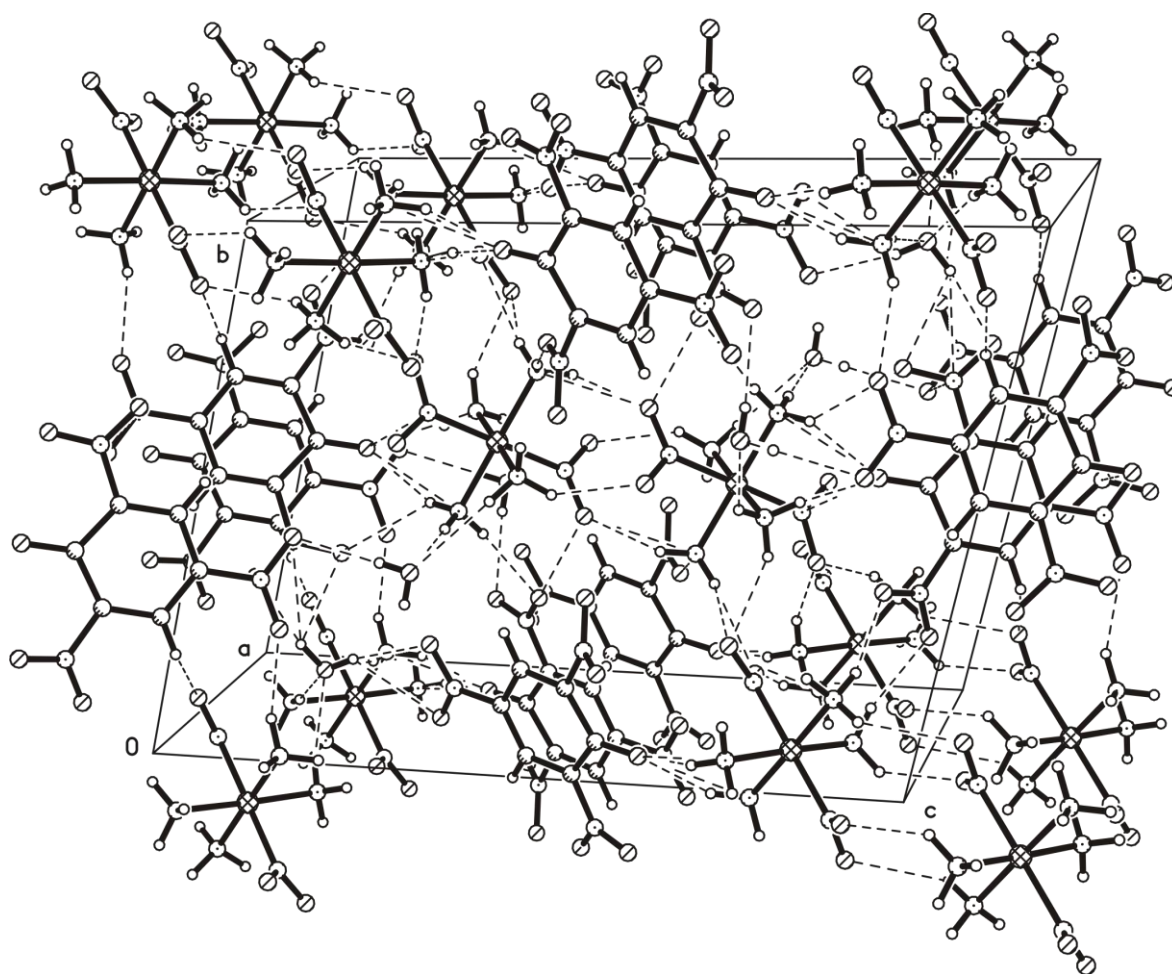


Fig. 18 Packing diagram of *trans*-[Co(NH₃)₄(NO₂)₂](picrate)·H₂O (**9**). View along the *a* axis.

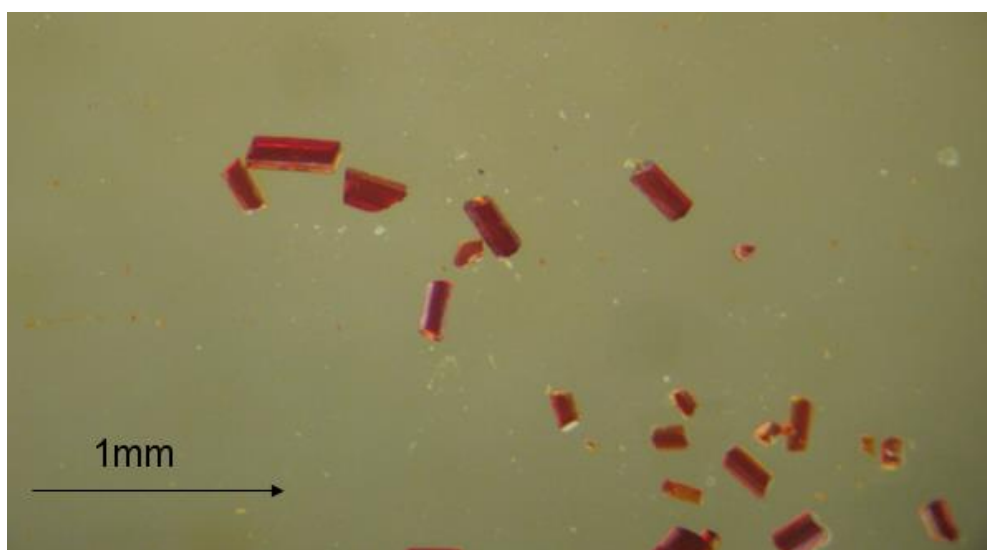


Fig. 19 Single-crystals of $[\text{Co}(\text{NH}_3)_5\text{N}_3](\text{picrate})_2$ (**10**).

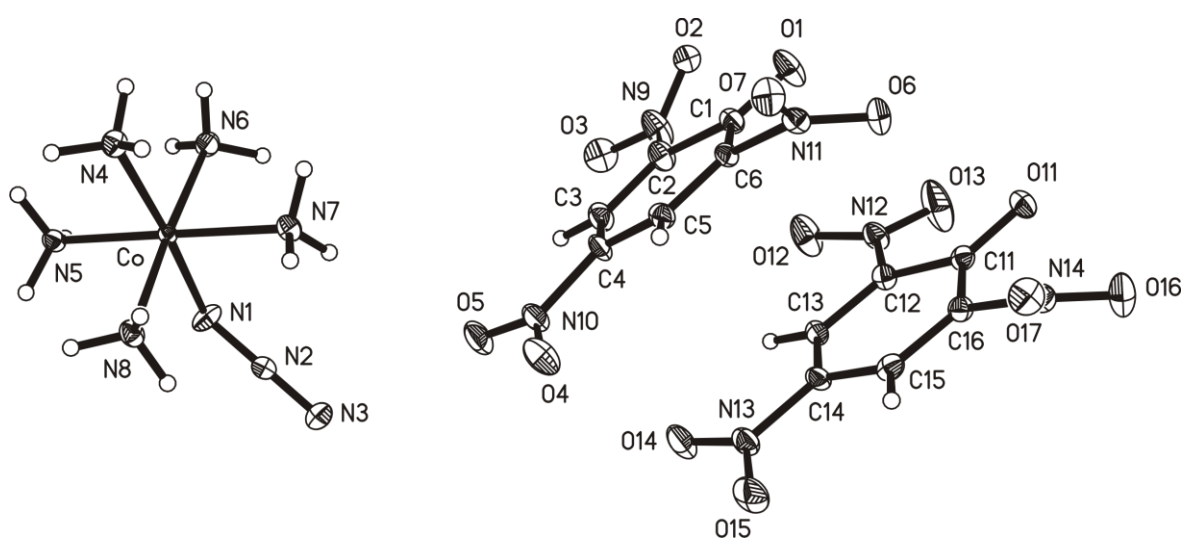


Fig. 20 ORTEP drawing of $[\text{Co}(\text{NH}_3)_5\text{N}_3](\text{picrate})_2$ (**10**) at the 50% probability level. Selected bond lengths [\AA] and angles [$^\circ$]: Co-N(5) 1.945(2), Co-N(1) 1.950(3), Co-N(8) 1.961(2), Co-N(4) 1.967(2), Co-N(6) 1.975(2), Co-N(7) 1.980(2), N(1)-Co-N(4) 175.08(9).

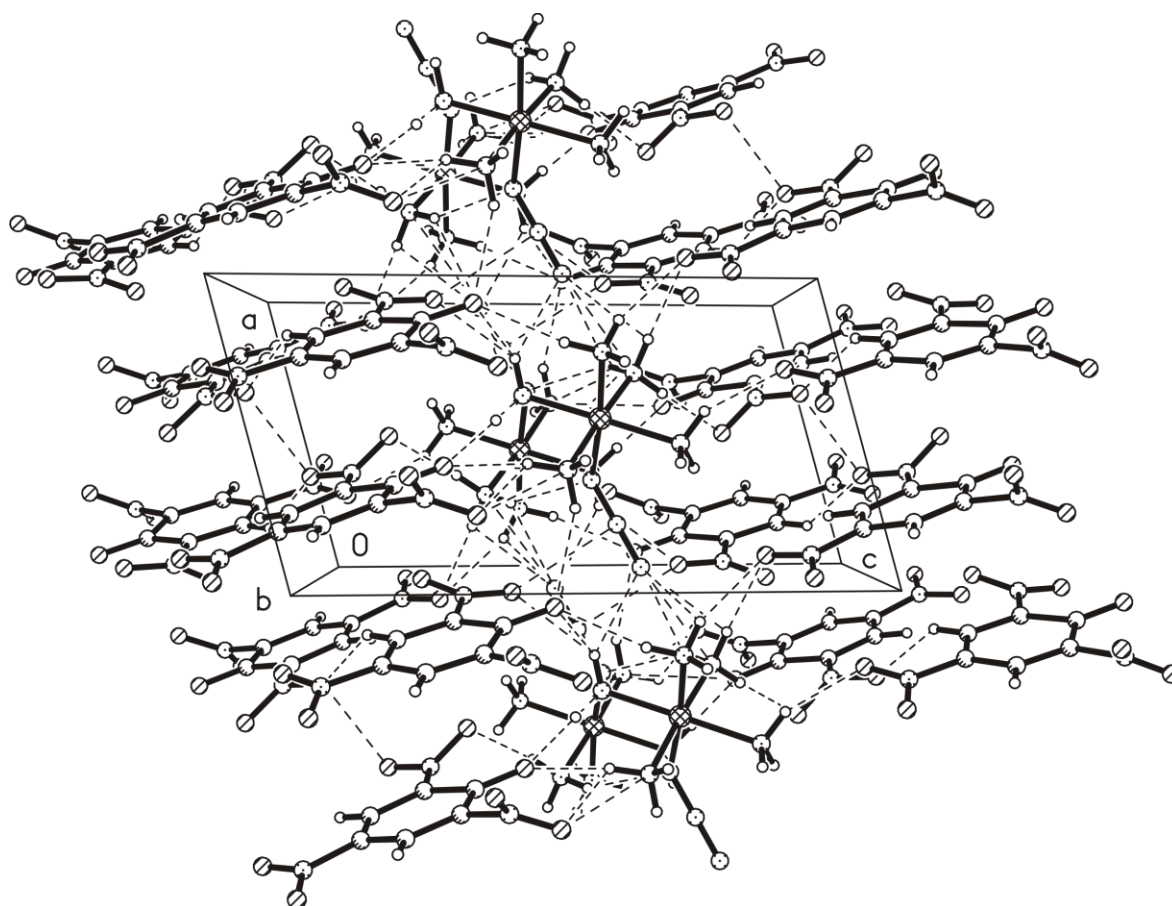


Fig. 21 Packing diagram of $[\text{Co}(\text{NH}_3)_5\text{N}_3](\text{picrate})_2$ (**10**). View along the b axis.

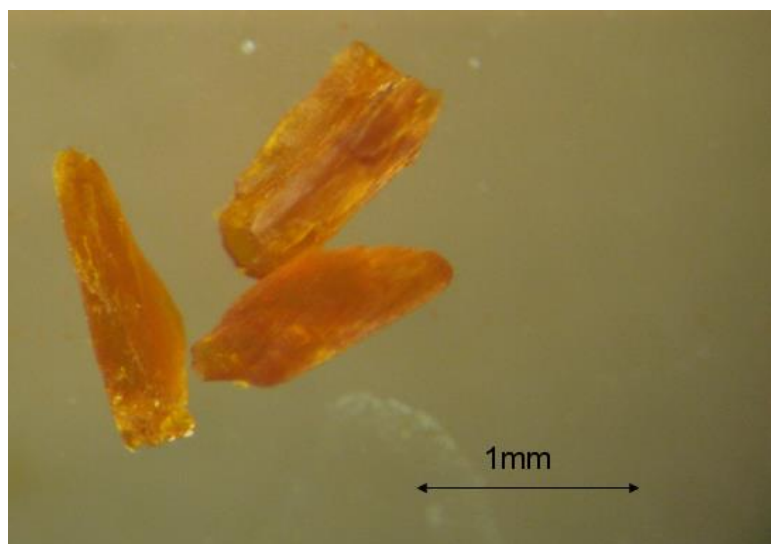


Fig. 22 Single-crystals of *trans*-[Co(NH₃)₄(NO₂)₂](DPA)·H₂O (**12**).

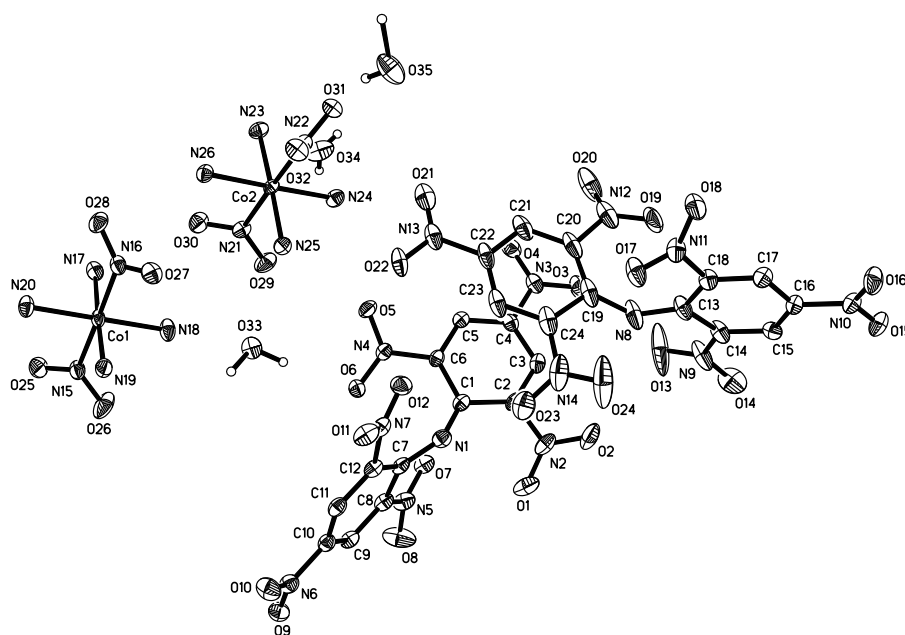


Fig. 23 ORTEP drawing of *trans*-[Co(NH₃)₄(NO₂)₂](DPA)·H₂O (**12**) at the 50% probability level. Selected bond lengths [Å] and angles [°]: Co(1)-N(16) 1.9298(18), Co(1)-N(15) 1.9353(18), Co(1)-N(19) 1.953(2), Co(1)-N(20) 1.9543(17), Co(1)-N(17) 1.9607(19), Co(1)-N(18) 1.9616(18), Co(2)-N(22) 1.9325(17), Co(2)-N(24) 1.9416(18), Co(2)-N(21) 1.9425(18), Co(2)-N(26) 1.9534(17), Co(2)-N(23) 1.9545(17), Co(2)-N(25) 1.9575(17), C(1)-N(1) 1.292(3), C(7)-N(1) 1.312(3), C(13)-N(8) 1.295(3), C(19)-N(8) 1.314(3), C(1)-N(1)-C(7) 144.45(19), C(13)-N(8)-C(19) 140.8(2).

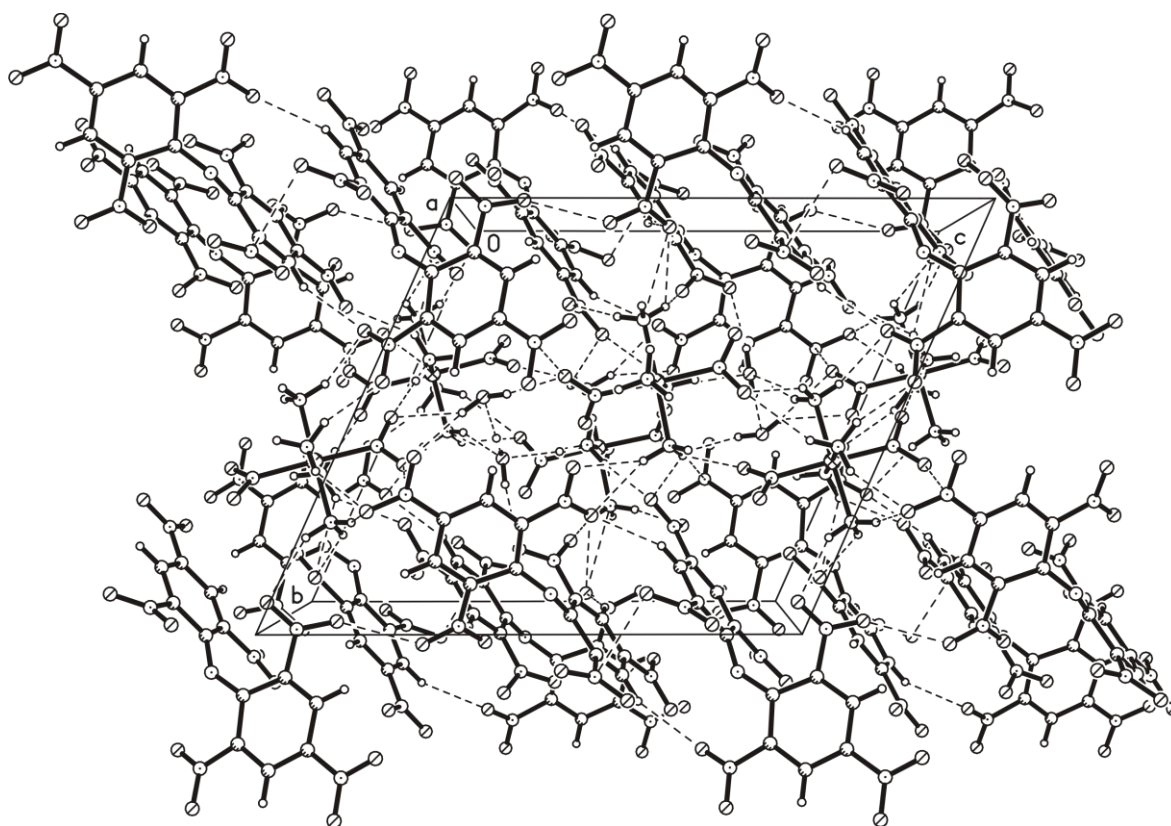


Fig. 24 Packing diagram of *trans*-[Co(NH₃)₄(NO₂)₂](DPA)·H₂O (**12**). View along the a axis.

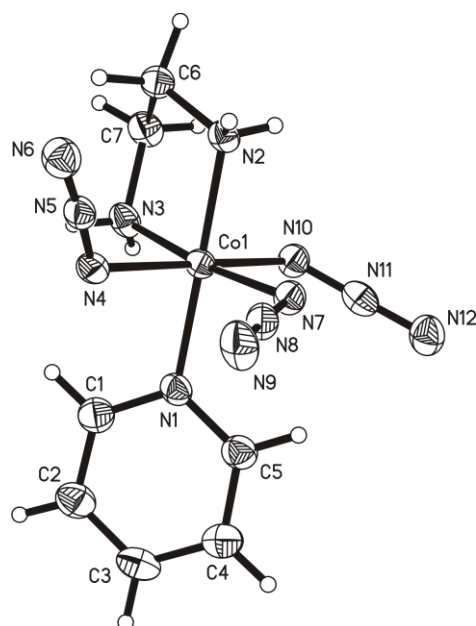


Fig. 25 ORTEP drawing of $[\text{Co}(\text{en})(\text{py})(\text{N}_3)_3]$ (**13**) at the 50% probability level. Selected bond lengths [\AA] and angles [$^\circ$]: $\text{Co}(1)\text{-N}(4)$ 1.946(2), $\text{Co}(1)\text{-N}(2)$ 1.948(2), $\text{Co}(1)\text{-N}(7)$ 1.962(2), $\text{Co}(1)\text{-N}(10)$ 1.969(2), $\text{Co}(1)\text{-N}(3)$ 1.973(2), $\text{Co}(1)\text{-N}(1)$ 1.975(2), $\text{N}(2)\text{-Co}(1)\text{-N}(1)$ 179.39(10), $\text{N}(4)\text{-Co}(1)\text{-N}(10)$ 174.59(10).

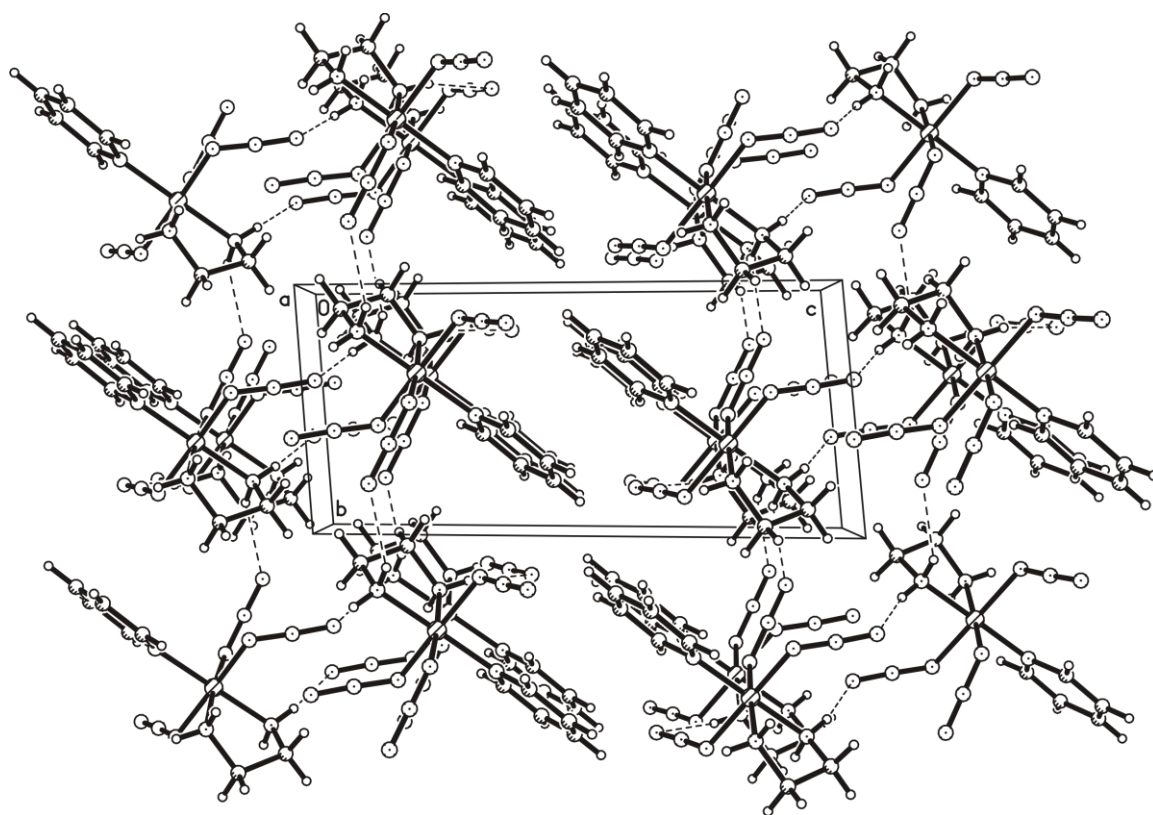


Fig. 26 Packing diagram of $\text{trans-}[\text{Co}(\text{NH}_3)_4(\text{NO}_2)_2](\text{DPA})\cdot\text{H}_2\text{O}$ (**12**). View along the *a* axis.



Fig. 27 BAM drophammer setup used in this study.



Fig. 28 BAM friction test setup used in this study.