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ARTICLE

Thermally cleavable Imine Base / Isocyanate Adducts and Oligomers suitable as Initiators for Radical Homo- and Copolymerization

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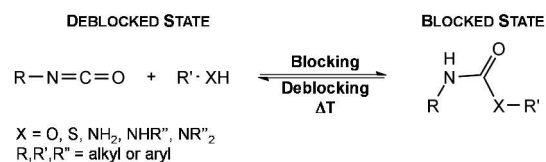
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The addition of isocyanates to C=N double bonds of imines gives triazindione heterocycle structures; their thermal properties are reported. Mono-isocyanates were used to form 2:1 adducts with the imine bases 1,8-diazabicyclo[5.4.0]undec-7-ene (DBU), 1,5-diazabicyclo[4.3.0]non-5-ene (DBN) and 2-*tert*-butyl-1,1,3,3-tetramethylguanidine (*t*BuTMG). A 2:1 stoichiometry of the adducts was proven by NMR and IR spectroscopy, and single crystal X-Ray diffraction; certain cleavage temperatures (70 and 160 °C) were measured. Thermal analysis (TG-MS) of adducts indicates the release of free isocyanate during adduct cleavage. Furthermore, a new class of step-growth oligomers ($M_N = 750\text{--}7,000 \text{ g}\cdot\text{mol}^{-1}$) composed of multi-functional isocyanates and these imine bases was introduced. Their systematic spectroscopic and thermal analysis is shown revealing the similarity in their chemical properties to the 2:1 adducts. Radical homo- and copolymerization of acrylates is initiated by the meta-stable adducts and oligomers of this work; the generation of novel telomeric block-copolymer architectures composed of polyacrylate and oligourea building blocks is demonstrated.

Introduction

Since their discovery in 1848 by WURTZ, isocyanates play an important role in industry.^[1] The synthesis of polyurethanes by the reaction of a polyester diol with a diisocyanate, discovered by BAYER and his co-workers in 1937,^[2] made diisocyanates to one of the major chemicals produced in the world.^[3,4] The main advantage of the use of isocyanates in industry is their high reactivity even at low temperatures or in high viscous systems ensuring a rapid conversion and an efficient mode of operation. One major drawback is the inherent toxicity of isocyanates.^[3] An approach to solve this problem is to temporarily transfer the isocyanate into meta-stable adducts, which are formed by reversible, isocyanate-based covalent bonds. The most common blocking agents bear hydrogen active nucleophile functionalities of type X-H, where X can be oxygen, sulfur, nitrogen or carbon. The reagents have a moderate acidity in the pKa range of 10 to 20 and facily undergo addition reactions with the electrophile isocyanate group at low temperatures, as depicted in Scheme 1.

At lower temperatures, isocyanates are in a blocked state. Therefore, the reactivity is decreased and hence the toxicity is remarkably reduced.^[3] Free isocyanates can be regenerated out of these adducts at elevated temperatures. The reversible blocking of isocyanates plays an important role in

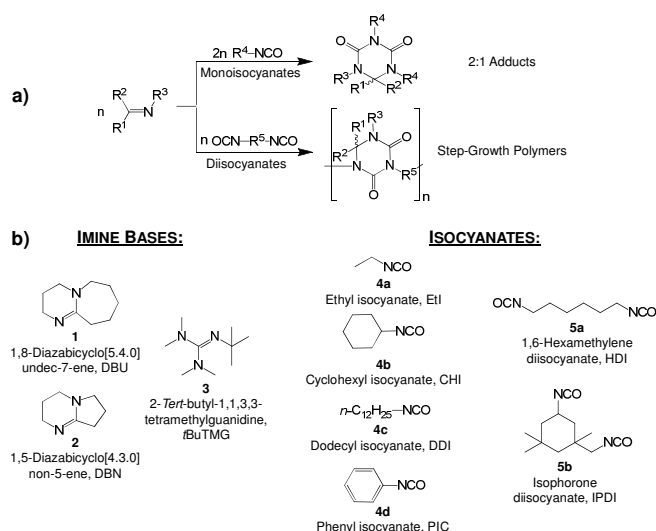


Scheme 1. Principle of established isocyanate blocking reactions with acidic X-H compounds.

polyurethane engineering processes which was reported extensively.^[3,5-8] Among others, malonic esters, oximes, phenols, lactams and secondary aniline derivatives have been studied in detail.^[9-14] Furthermore, the formation of urethdions, dimers of isocyanate groups, is another way to produce blocked isocyanates. The advantage of those systems is that no additional blocking agent is required. However, this method is solely applicable to aromatic isocyanates. In addition, specific catalysts such as toxic phosphines are necessary for their formation and, due the reversibility of the reaction, residues of highly reactive NCO-groups still are found in the reaction mixture.^[15] A recent contributions^[16] to the field revealed a new method for introducing an isocyanate group, blocked by 3,5-pyrazol, into a polymer architectures. However, post-modification of the isocyanate group by amines and alcohols was demonstrated to occurred at 130 °C. Thus, up to nowadays it is highly desirable in industry and academic research to find

methods for the selective control of the isocyanate reactivity for a reasonable and safe handling of this versatile group.

In this work, we report on the reversible adduct formation of mono- and bifunctional isocyanates with imine bases. Thermal properties and a detailed structure analysis of the products are provided. The adduct formation mechanism of mono-functional isocyanates with amidines was investigated earlier by ULRICH, RICHTER and co-workers.^[17–19] They reported that their strong exothermic reaction at room temperature results into the formation of 2:1 adducts bearing six-membered heterocyclic rings, as shown in Scheme 2a.^[4,20]



Scheme 2. (a) Schematic view of the addition reactions of mono- and diisocyanates and imines yielding 2:1 adducts and step-growth polymers as well as (b) structures of imine bases and isocyanates used in this study. R^1 – R^5 are alkyl groups.

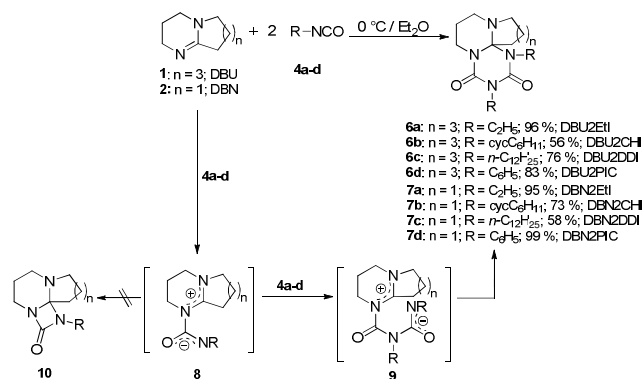
Recently, we have shown that isocyanates undergo radical formation when treated with imine bases and vinyl monomers.^[21,22] Subject of the present study is the investigation on obvious adduct formations of amidine and guanidine derivatives with isocyanates as special cases of this investigation. The specific novelty of this work is that we employed a non-acidic blocking strategy. Furthermore, the quasi-bi-functionality of the imine moiety suggests its potential as a component for step-growth polymerization with diisocyanates. Herein, we show that multiple reactions of diisocyanates and imine bases result in a polyaddition mechanism yielding meta-stable step-growth polymers with a cyclic repeating unit (see Scheme 2a). A systematic study of this new polymer class is provided.

The structures of reactants used in this work are given in Scheme 2b. Thermogravimetric analysis (TGA), differential scanning calorimetry (DSC), and mass spectrometry are used to clarify the cleavage process in detail. With the approach described in this work we generate blocked isocyanates that are stable at room temperature and release free isocyanate at slightly elevated temperatures in a range between 70 and 160 °C. The adduct cleavage temperature is mainly influenced by the isocyanate reactivity and steric hinderance; adducts and oligomers composed of aromatic or voluminous isocyanates have a lower cleavage temperature than aliphatic isocyanates.

The cleavage analysis results reveal that radical intermediates are formed. In this context, we investigated the potential of the adducts and oligomers as initiators for the radical homo- and copolymerization of acrylic vinyl monomers. We showed that, during the cleavage of the appropriate imine base–isocyanate adducts and oligomers, radical species are formed that, in combination with acrylates, result in the formation of high molecular weight polyacrylates. When using oligomers for the acrylate polymerization, novel telomeric block-copolymers composed of polyacrylate and oligourea building blocks are formed that are characterized by gel permeation chromatography (GPC) and nuclear magnetic resonance (NMR) spectroscopy. With our approach, a novel path for the reversible formation of blocked isocyanates is given that enables the fabrication of functional block copolymers and materials based on the isocyanate structure.

Results and discussion

The reaction of isocyanates with imines proceeds immediately by the formation of 2:1 adducts at room temperature, as indicated in the experimental section. This addition is similar to most reactions of isocyanates with *N*-nucleophiles an exothermic reaction.^[4,20] Hence, all syntheses are carried out at 0 °C to prevent both side product formation and the back reaction. The DBN and DBU adducts with alkyl and aryl mono isocyanates are obtained in high to almost quantitative yields (56–99 %). The adduct formation mechanism was proposed earlier by ULRICH, RICHTER and co-workers;^[17–19] the suggested reaction path and analysis results of this work are given in Scheme 3.



Scheme 3. Schematic formation of the adducts 6 and 7 and the mechanism of stepwise addition of two mol isocyanate to one mol imine.^[17–19]

Adduct formation proceeds *via* a two-step mechanism. In the initial step, one mole of isocyanate reacts with the nucleophilic imine function of the amidine to the zwitterionic intermediate 8. The rate of this reaction is mainly determined by the electron withdrawing character and the steric hinderance of the group *R* in 4 (see Scheme 2). The reaction continues by addition of another mole of isocyanate to the zwitterionic intermediate 8 to generate 9. The second isocyanate addition occurs faster than the first one. The formation of tricycles 10 is not observed; moreover, the addition of another mole of isocyanate yields tricycles 6 and 7. The structure of the so-formed six-membered rings of adducts 6 and 7 consists of a triazindione heterocycle. In Figure 1, the results of the X-ray single crystal structure of

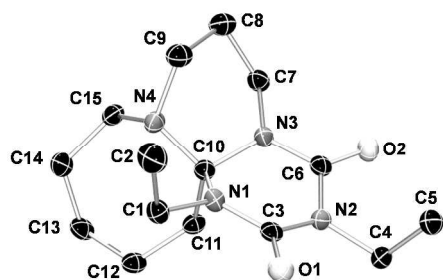


Figure 1. Results of the single-crystal X-ray diffraction analysis of **6a** (97.3 % ellipsoid probability). **6a** is formed racemic, the S-enantiomer is depicted. A table of relevant bond lengths and angles is shown in the SI.

6a is shown. The analysis clearly manifests the hypothesis of the formation of a tricyclic structure containing a triazindione heterocycle. In particular, the quaternary carbon C10 (Fig. 1) is prominent for structures **6** and **7** that give characteristic ^{13}C NMR shifts. The presence of tricyclic structures **6** and **7** is also proven by the signal of the quaternary C10 in the ^{13}C NMR spectrum close to 90 ppm (CDCl_3); more detailed information are included in the Supporting Information Section.

The adduct formation using *t*BuTMG, in conjunction with isocyanates is more complex than that of **6** and **7**. Analysis results, single crystal structure of a model compound, as well as the mechanistic background of this reaction and the cleavage products, are provided in the Supporting Information (Fig. S2). To elucidate the thermal profiles and the processes occurring at elevated temperatures of compounds **6** and **7**, the DSC curve of **7a** is shown in Figure 2.

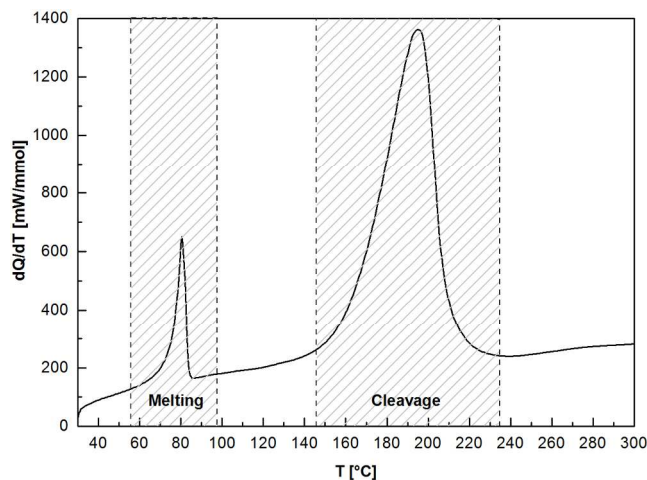


Figure 2. DSC-curve of **7a** (DBN2EtI) at a heating rate of 10 K/min. The melting occurs at $T_{\text{onset}} = 75.8$ °C, $T_{\text{peak}} = 80.3$ °C, $Q = 0.27$ $\text{J}\cdot\text{mol}^{-1}$ and cleavage is detected at $T_{\text{onset}} = 161.8$ °C, $T_{\text{peak}} = 194.5$ °C, $Q = 2.87$ $\text{J}\cdot\text{mol}^{-1}$.

When systematically heating up the sample to 300 °C, the thermogram shows a sequence of two endothermic processes. The first peak ($T_{\text{onset}} = 75.8$ °C, $T_{\text{peak}} = 80.3$ °C, $Q = 0.27$ $\text{J}\cdot\text{mol}^{-1}$) indicates the melting of **7a**, while at higher temperatures ($T_{\text{onset}} = 161.8$ °C, $T_{\text{peak}} = 194.5$ °C, $Q = 2.87$ $\text{J}\cdot\text{mol}^{-1}$), cleavage of **7a** starts which is the opposite reaction to the exothermic **7a** formation. Cleavage temperatures of **6** and **7** are in the range

between 69 and 162 °C; the corresponding cleavage enthalpies are 0.02–2.87 $\text{J}\cdot\text{mol}^{-1}$ and a function of the isocyanate type. Focusing solely on the isocyanate, the cleavage enthalpies and temperatures, and hence the stability, decreases for the following: $\text{R} = \text{C}_2\text{H}_5 > \text{cyc-C}_6\text{H}_{11} > n\text{-C}_{12}\text{H}_{25} > \text{C}_6\text{H}_5$. There are no significant adduct cleavage temperature differences when changing the imine base type (DBN or DBU).

DSC curves of **6**, **7** and **11** are measured to verify their thermal properties; analysis details of the adducts are given in Table 1.

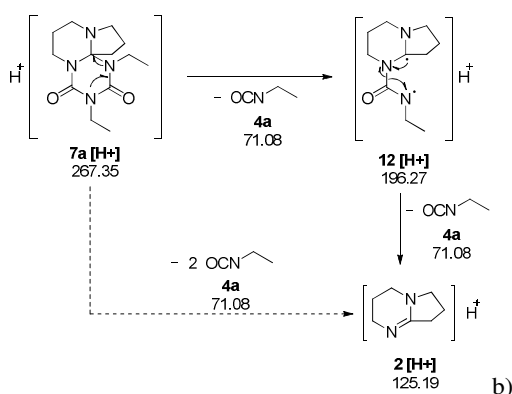
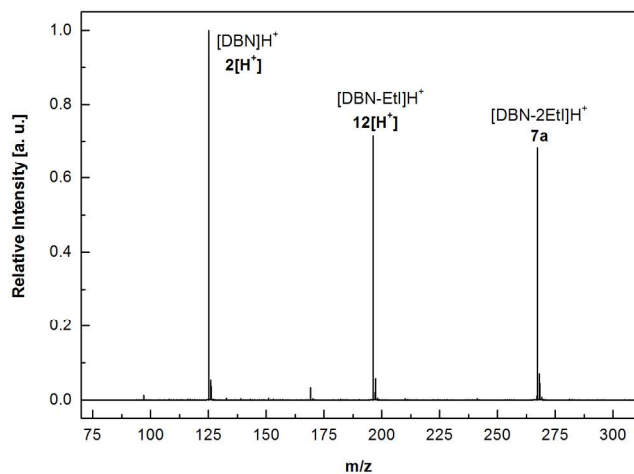
Table 1. Thermal analysis and synthesis details of adducts **6**, **7** and **11**.

Adduct	R =	T_{cleavage} [°C]	Q_{cleavage} [$\text{J}\cdot\text{mol}^{-1}$]
6a (DBU2EtI)	C_2H_5	154	1.219
6b (DBU2CHI)	$\text{cyc-C}_6\text{H}_{11}$	115	0.580
6b (DBU2DDI)	$\text{C}_{12}\text{H}_{25}$	76	0.035
6b (DBU2PIC)	C_6H_5	69	0.126
7a (DBN2EtI)	C_2H_5	162	2.866
7b (DBN2CHI)	$\text{cyc-C}_6\text{H}_{11}$	125	0.623
7c (DBN2DDI)	$\text{C}_{12}\text{H}_{25}$	101	0.241
7d (DBN2PIC)	C_6H_5	73	0.019
11 (TMG3EtI) ^[a]	C_2H_5	176	1.586

[a] More details on the formation mechanism of **11** are included in the Supporting Information.

It was found that the adduct cleavage temperatures are much lower compared to established isocyanate blocking systems, such as blocked isocyanates based on ϵ -caprolactam, phenols or triazoles.^[5–8,11–13] Monomeric adducts are obtained by the reaction of DBU, DBN and *t*BuTMG with the monoisocyanates ethyl isocyanate **4a** (EtI), cyclohexyl isocyanate **4b** (CHI), dodecyl isocyanate **4c** (DDI) and phenyl isocyanate **4d** (PIC).

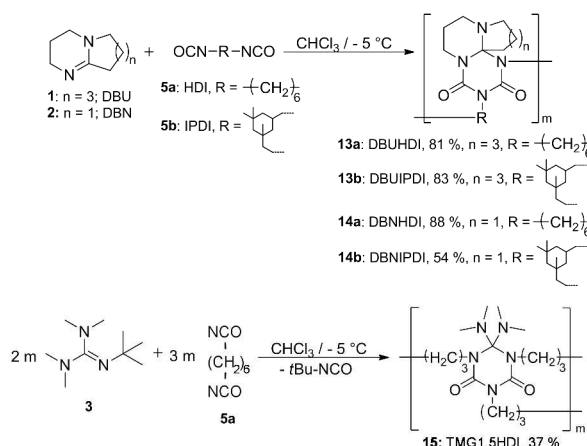
We recorded ESI-MS and TG-MS spectra to further study separate fragments from the cleavage. The ESI-MS spectrum of **7a** is shown in Figure 3 and a reasonable fragmentation mechanism is given in Figure 3a. Fragments at $m/z = 196$ ($[\text{DBU-EtI}]^+$) and $m/z = 125$ ($[\text{DBU}]^+$) are clearly detected. This indicates a stepwise elimination of two moles isocyanate from **7a**. The intermediate **12** is assumed to have a radical character. The release of free isocyanate from **7a** at elevated temperatures is also confirmed by TG-MS measurements; spectra are in the SI. In the TG-MS analysis of **7a**, EtI **4a** is detected at 192.5 °C. The free isocyanate **4a** can be trapped by nucleophilic substances such as methanol or *n*-butylamine, quantitatively.^[23,24] When using diisocyanates instead of monoisocyanates, the reaction proceeds as a polyaddition reaction to form oligomeric structures with the triazindione repeating unit as shown in Scheme 4. Oligomers **13–15** were obtained after washing and drying as foamy white solids that are soluble in polar organic solvents such as chloroform and methanol. Number average molecular weights of **13–15** are in a



a)

b)

Figure 3. ESI-MS-spectrum of DBN2EtI **7a** (a) and presumed fragmentation mechanism of **7a** (b). Fragments in a) correspond to species: $m/z = 125.11$ [DBN] H^+ , $m/z = 196.15$ [DBN-EtI] H^+ and $m/z = 267.21$ [DBN2EtI] H^+ . The excitation energy is 8,500 V.



Scheme 4. Schematic view of the oligomer synthesis DBUHDI **13a**, DBUIPDI **13b**, DBNHDI **14a**, DBNIPDI **14b** and TMG1.5HDI **15**.

range between 750 to 7,000g/mol (poly dispersity = $M_w/M_n = 1.35$ – 1.75 , Table 2). The degree of polymerization is between 2 and 21, so we consider the products to be oligomers. Moreover, oligomers of *t*BuTMG and isophorone diisocyanate (IPDI) are

not obtained; a drastic steric hindrance of both monomers prevents any

Table 2. Imine base / Isocyanate oligomers and GPC analysis results according to Scheme 4.

Amine	Isocyanate	Oligomer	M_N [$g \cdot mol^{-1}$]	M_w/M_n
		13a	5,000	1.71
		13b	1,300	1.35
		14a	2,400	1.60
		14b	6,400	1.48
		15	780	1.30

reaction. The presence of the cyclic triazindione repeating units in oligomers **13–15** is readily identified by ^{13}C NMR spectroscopy on the quaternary C-atom that is formed during the cyclization reaction. The ^{13}C NMR spectrum shows the characteristic quaternary C-atom of the oligomers resonating at 92 (**13**, **14**) and 103 ppm (**15**) (Figure 1). Furthermore, a broad signal at 121 ppm in the ^{13}C NMR spectra of **13–15** reveals the presence of isocyanate end groups. This hypothesis is also confirmed by IR spectroscopy. Vibrational spectra of oligomers **13–15** show a broad band at 2170 cm^{-1} that can be assigned to the asymmetric stretching mode of the NCO group, ν_{as} . Further analysis results of oligomers **13–15** are summarized in the Supporting Information. A more detailed synthetic procedure as well as further chemical and physical properties of those oligomers obtainable from amidines and diisocyanates will be reported in a separate work.

The thermal properties of compounds **13–15** are similar to their monomeric adducts. Figure 4 shows the cyclic DSC analysis of **14a**.

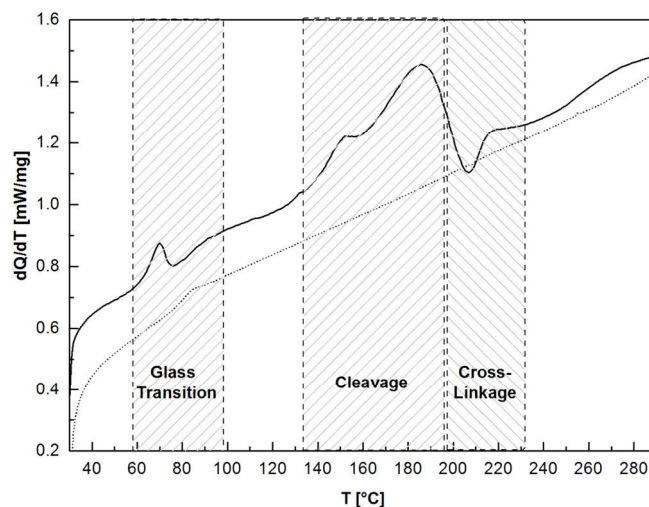


Figure 4. Cyclic DSC curves of DBNHDI **14a** (1st –, 2nd ...) at a heating rate of 10 K/min. At the first thermal treatment the glass-transition occurs at $T_{onset} = 62.7\text{ °C}$, $T_{peak} = 70.0\text{ °C}$, $Q = 4.50\text{ J} \cdot g^{-1}$, cleavage at $T_{onset} = 150.1\text{ °C}$, $T_{peak} = 184.8\text{ °C}$, $Q = 64.95\text{ J} \cdot g^{-1}$ and a cross-linking process is detected at $T_{onset} = 199.7\text{ °C}$, $T_{peak} = 207.0\text{ °C}$, $Q = -7.03\text{ J} \cdot g^{-1}$.

¹. A glass transition is detected at the second heating at $T_{\text{peak}} = 84$ °C.

When heating **14a** initially, a glass transition process occurring at 63 °C ($Q = 4.50 \text{ J}\cdot\text{g}^{-1}$) was recorded. The reversibility of this process was proven by testing the solubility of the oligomer. A sample heated to 70 °C for 2 h is still completely soluble in chloroform. This glass transition is followed by an endothermic process at 150 °C ($Q = 64.95 \text{ J}\cdot\text{g}^{-1}$) that can be assigned to the engaging cleavage. Systematically, after heating a sample to 185 °C for a time period of 3 h, the foamy oligomer becomes insoluble in chloroform. After heating the sample above the cleavage, an exothermic process is detected ($T_{\text{onset}} \sim 200$ °C, $Q = -7.03 \text{ J}\cdot\text{g}^{-1}$) that can be assigned to an irreversible transformation and cross-linkage of **14a**. An insoluble brownish rubber-like product was obtained after the thermal treatment indicating an inter- and intramolecular cross-linkage. Free isocyanate functions cannot be detected in the product, which was proven by IR spectroscopy. After the second thermal treatment, no cleavage and cross-linking processes are detected; only a glass transition (84 °C) of the tempered oligomer was recorded. In Figure S3 in the Supporting Information liquid high resolution ^{13}C NMR and solid state ^{13}C CP MAS NMR spectra of tempered **14a** at 70, 185 and 210 °C are presented, supporting the observations described above. In addition, IR studies on the cleavage of adduct **7d** are shown in Figure S5. It is notable that products containing an aromatic or sterically hindered isocyanate show the lowest cleavage temperature, which is caused by the increased structural stress in these compounds. In Table 3 all adduct and oligomer cleavage temperatures presented in this work are summarized.

Table 3. Comparison of adduct and oligomer cleavage temperatures (°C) of this work.

Isocyanate	Imine	DBU	DBN	<i>t</i> BuTMG
PIC		69	72	–
DDI	$n\text{-C}_{12}\text{H}_{25}\text{-NCO}$	76	73	–
CHI		115	125	–
EtI		154	162	176
IPDI		116	137	–
HDI		130	106	152

One basic question still remains – what is the nature of the fragments at the adduct and oligomer cleavage? For the recyclization mechanism at thermal treatment, we assumed the formation of radical species (Figure 3). To prove our assumption the adducts are thermally cleaved in the presence of acrylic monomers. Remarkably, the monomeric and oligomeric imine base–isocyanate adducts initiate a radical polymerization of methyl methacrylate (MMA) and other acrylic monomers simply by thermal treatment, resulting in high molecular weight polymers. To evidently show the radical character of the propagating species, copolymerization experiments were carried out using co-monomers MMA and methacrylonitrile (MAN). We determined the copolymerization parameters, r_{MMA} and r_{MAN} , by standard methods^[26,27] and compared the results with literature known r -values of anionic and radical

copolymerization of these co-monomers. Table 4 shows the experimental results of the copolymerization using DBU2EtI **6a** as the initiator at 80 °C and results from literature, according to radical and anionic copolymerization experiments.

Table 4. Copolymerization parameters r_{MMA} , r_{MAN} , and azeotrope values of the copolymerization of MMA and MAN as well as previously reported values of a radical and anionic polymerization of these co-monomers.

Initiation	r_{MMA}	r_{MAN}	Azeotrope Value
DBU2EtI 6a	0.50	1.40	–
Radically (AIBN) ^[26]	0.67	0.65	0.48
Anionic (<i>n</i> -butyl lithium) ^[26]	0.67	5.20	–

These results clearly confirm that the propagating species of this polymerization is of radical nature.

Furthermore, imine base / isocyanate adducts and oligomers were used for a radical polymerization of MMA. Polymerization conditions and MMA conversions X_{MMA} as well as GPC analysis results of PMMA of selected MMA-polymerizations using adducts **6a,b,d**, **7** and **11**, as well as oligomers **13** and **15** are shown in Table 5.

Table 5. Results of MMA polymerization and analysis results using imine base–isocyanate adducts **6a,b,d**, **7**, **11** and oligomers **13** and **15** as initiators. All polymerizations were carried out at 80 °C. X_{MMA} is the monomer conversion.

Adduct (Ad)	[MMA] / M	[Adduct] / M	t / h	X_{MMA} / %	$M_{\text{N}} / \text{g}\cdot\text{mol}^{-1}$	$M_{\text{w}}/M_{\text{n}}$
6a	3.53	0.49	6	13	368.400	1.66
6b	3.52	0.48	18	8	33.000	1.35
6d	3.52	0.46	2	74	36.700	1.72
13a	9.09	0.16	2	13	354.800	1.85
13b	5.22	0.08	4	21	112.200	2.38
7a	3.59	0.49	6	15	264.300	1.41
7b	3.52	0.48	18	25	33.100	2.26
7c	3.52	0.48	18	7	82.400	1.62
7d	3.52	0.48	2	82	23.500	1.89
11	3.51	0.52	6	55	112.500	1.91
15	9.98	0.06	1	28	180.100	2.26

The higher the stability of adducts and oligomers **6a,b,d**, **7**, **11** and **13**, **15**, the higher the average molecular weight of PMMA is, which is consistent with the corresponding initiator concentration. As an example, for the least stable adducts, DBU2PIC **6d** and DBN2PIC **7d**, the MMA polymerization proceeds rapidly until all of the initiator is consumed. Polymerization of MMA using **6d** and **7d** is possible even at moderate temperatures (45–70 °C). Consequently, low molecular weight PMMA is isolated ($M_{\text{N}} = 23.500\text{--}36.700 \text{ g/mol}$). In comparison, the more stable adducts with higher

cleavage temperatures give high molecular weight PMMA (112.000–368.400 g/mol) and non-cleaved adduct is isolated after polymerization. However, in the case of the polymerization with **6b** whose cleavage temperature (115 °C) is below that of **6a** (154 °C), the polymer yield is significantly lower. The authors assume that the stability of the radicals, formed at the cleavage of **6b** (secondary alkyl isocyanate) is increased disfavoring an initiation. We are able to reuse stable adducts for polymerization after their purification. Adducts and oligomers are isolated from the reaction mixture by precipitating the product PMMA in methanol. The initiators are completely soluble in this solvent.

Interestingly, when polymerizing MMA using oligomers **13** and **14**, even after purification of the product still **13** and **14** is being detected in the ^1H , ^{13}C NMR and GPC. Product NMR spectra and GPC graphs resulting from the MMA polymerization using DBUHDI **13a** are depicted in Figure 5. No starting material signal is found in the GPC of PMMA-*b*-DBUHDI, indicating the presence of only one product species. The polymerization product is – like PMMA – insoluble in methanol. In contrast, DBUHDI **13a** is well soluble in methanol. Even after an extensive purification of the methanol insoluble product by multiple precipitation and washing in methanol, still signals of **13a** were still detected. Obviously, the oligomers are covalently

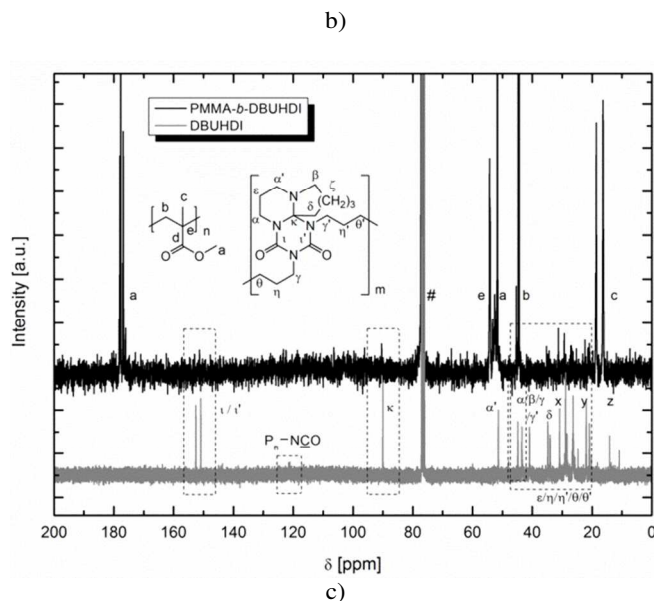
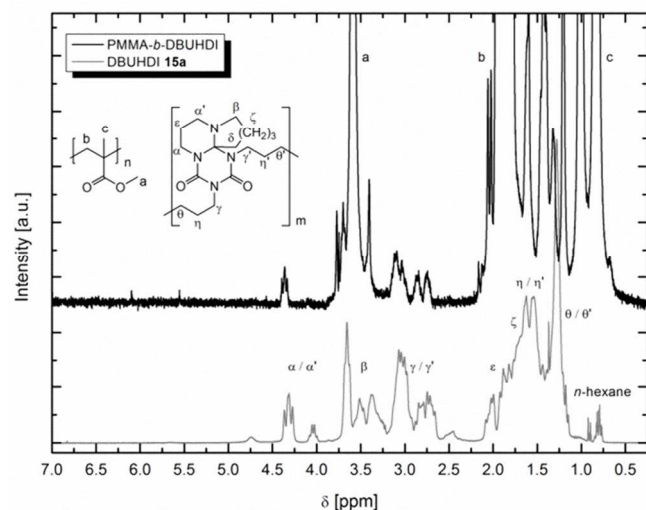
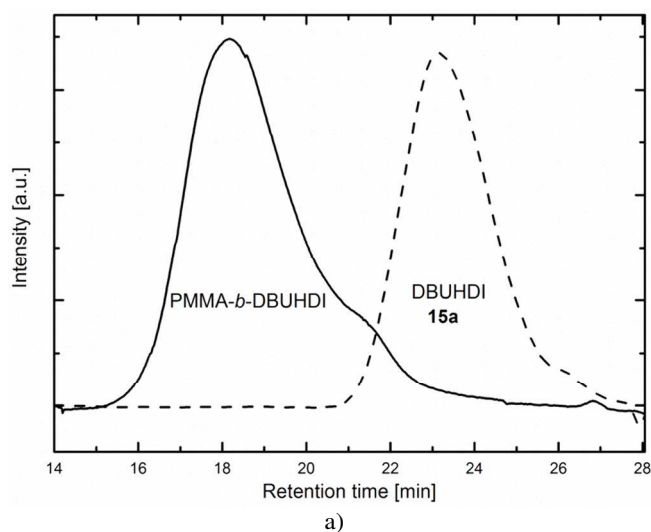


Figure 5. (a) GPC, (b) ^1H - and (c) ^{13}C NMR analysis of **13a** (–) and block-copolymer PMMA-*b*-DBUHDI (–). Signals of DBUHDI in the ^{13}C NMR spectrum of PMMA-*b*-DBUHDI are highlighted by boxes. The number average molecular weight M_N of **13a** (–) is 5.000 g/mol (PDI = 1.71) and that of PMMA-*b*-DBUHDI (–) is 354.800 g/mol (PDI = 1.85).

bonded to the PMMA chain and a novel class of block-copolymers containing both a polyacrylate and an oligoureia backbone is generated. The AB-block structure of PMMA-*b*-DBUHDI is confirmed by the NMR spectra that have only one broad resonance of the OCH_3 ester group of PMMA. Taking into account quantitative elementary analysis and quantitative ^1H NMR analysis of the product, the PMMA:DBUHDI ratio is 52:1. Thus, short oligomer blocks of DBUHDI are covalently bonded to long PMMA chains. Hence, the structure more resembles that of a telomeric block-copolymer.

The short oligomer block is still intact in its reactivity. We record endothermic and exothermic peaks of DBUHDI in the telomeric block copolymers PMMA-*b*-DBUHDI and PMA-*b*-DBUHDI in the cyclic DSC, as shown in Figure 6.

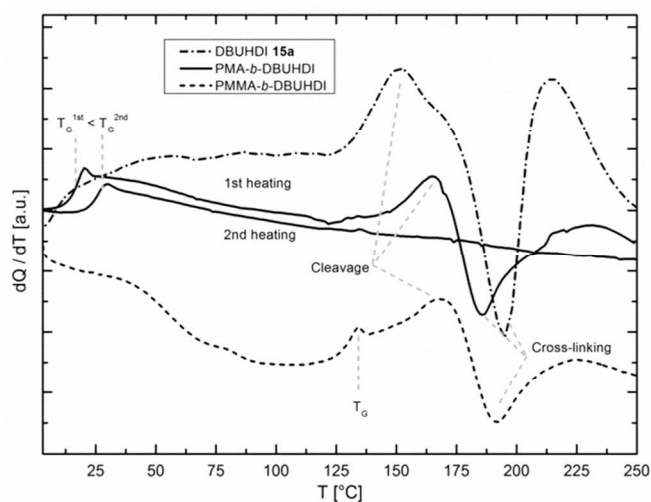


Figure 6. Cyclic DSC curves of DBUHDI (–), PMA-*b*-DBUHDI (1st –

, 2nd –) and PMMA-*b*-DBUHDI (---) at a heating rate of 10K·min⁻¹. DBUHDI: cleavage: $T_{\text{onset}} = 130\text{ }^{\circ}\text{C}$, $Q = 39.4\text{ J/g}$; cross-linkage: $T_{\text{onset}} = 180.0\text{ }^{\circ}\text{C}$, $Q = -38.5\text{ J/g}$; 2nd endothermic peak: $T_{\text{onset}} = 203.6\text{ }^{\circ}\text{C}$, $Q = 38.2\text{ J/g}$. Block copolymer 1st thermal treatment: PMMA-*b*-DBUHDI: glass transition $T_{\text{onset}} = 127\text{ }^{\circ}\text{C}$, endothermic peak $T_{\text{onset}} = 155\text{ }^{\circ}\text{C}$, exothermic peak $T_{\text{onset}} = 175\text{ }^{\circ}\text{C}$, endothermic peak $T_{\text{onset}} = 201\text{ }^{\circ}\text{C}$; PMA-*b*-DBUHDI: glass transition $T_{\text{onset}} = 13\text{ }^{\circ}\text{C}$, endothermic peak $T_{\text{onset}} = 144\text{ }^{\circ}\text{C}$, exothermic peak $T_{\text{onset}} = 168\text{ }^{\circ}\text{C}$, endothermic peak $T_{\text{onset}} = 200\text{ }^{\circ}\text{C}$. Block copolymer 2nd heating PMA-*b*-DBUHDI: glass transition $T_{\text{onset}} = 20\text{ }^{\circ}\text{C}$.

The characteristic DSC signals of DBUHDI are slightly shifted to increased temperatures in the block copolymers (Figure 6). After the thermal treatment of the block-copolymers, insoluble rubber-like (PMA-*b*-DBUHDI) and robust glassy (PMA-*b*-DBUHDI) polymers are obtained. Multiple cross-linking of the block-copolymer sequence caused by the radical degradation of the DBUHDI-block occurred. This process is irreversible and influences the properties of the polyacrylate block: no further peaks of DBUHDI are detected in PMA-*b*-DBUHDI at the second heating process. In addition, the glass transition temperature of PMA is shifted to higher temperatures (20 °C), indicating an increased rigidification of the polyacrylate backbone. These experiments reveal the novel block-copolymers to be promising candidates for the development of new types of programmable coatings and resins based on the acrylate and polyurea structure increment.

Conclusions

We have shown that the reaction of isocyanates and imines yields metastable thermal cleavable 2:1 adducts with a heterocyclic triazindione structure. Their stability is mainly a function of the isocyanate nature, which was proven by DSC measurements. The cleavage of the adducts occurs at low temperatures between 70 and 160 °C. Thermal treatment of the adducts leads to a stepwise re-cyclization process with the release of free isocyanate which is shown by ESI-MS studies. Thus, the use of the adducts as latent isocyanate sources is indicated. When converting diisocyanates with the imine bases of this work, a new class of metastable oligomers with six-membered triazindione rings as repeating units was formed. The average number molecular weights M_N of this step-growth oligomers is in a range between 750 and 7,000 g/mol. The oligomers bear reactive NCO-end groups and their cleavage temperatures are similar to the ones of the monomeric adducts. After the thermal treatment of the oligomers at elevated temperatures, insoluble rubber-like products are formed that result from the stepwise fragmentation of the instable tricyclic structures, followed by intra- and intermolecular rearrangements and cross-linkage. That adducts and oligomers were effectively used as initiator for radical polymerizations of acrylate monomers. Telomeric polyacrylate-oligourea block-copolymers structures were formed when using oligomers as initiators for the acrylate polymerization. These reactive, well-defined block-copolymers are a novel polymer class with a direct covalent bonding of polyacrylate chains to meta-stable oligourea based cyclic triazindione structures, which open doors for approaches for the synthesis of programmable hybrid coatings and resins based on these polymer classes.

Experimental section

Materials

The isocyanates ethyl isocyanate (EtI, Aldrich), phenyl isocyanate (PIC, ABCR), cyclohexyl isocyanate (CHI, Merck) dodecyl isocyanate (DDI, ABCR), 1,6-hexamethylene diisocyanate (HDI) and isophorone diisocyanate (IPDI) (Evonik Industries AG) and imine bases 1,8-diazabicyclo[5.4.0]undec-7-ene (DBU), 1,5-diazabicyclo[4.3.0]non-5-ene (DBN, Aldrich), and 2-*tert*-butyl-1,1,3,3-tetramethylguanidine (*t*BuTMG) were used as received. Monomers methyl methacrylate (MMA) and methacrylonitrile (MAN) were purchased from Sigma Aldrich, distilled prior to use and stored under Argon in the freezer (–20 °C). Diethyl ether (Et₂O) and chloroform were purified by standard methods.^[24]

Measurements

NMR spectra were recorded with a Bruker Avance DRX 250 apparatus (¹H: 250.13 MHz, ¹³C: 62.90 MHz) using CDCl₃ as solvent.

Size exclusion chromatography (SEC) measurements are performed using a PL-GPC 50 plus from Polymer Laboratories equipped with a PL-AS RT auto sampler and a PC-RI detector. A PLgel MIXED-D column is used with THF as eluent with a flow rate of 1 mL·min⁻¹ at 40 °C.

ATR-FT-MIR spectra were measured with a FTS-165 spectrometer from BioRad equipped with a golden gate device. Elemental analysis was performed with a Vario El apparatus from Elementar Analysensysteme GmbH Hanau.

The ESI mass spectra were recorded with a micrOTOF QII (ESI-Qq-TOF) from Bruker Daltonik. Excitation voltages range between 4,500 and 8,500 V.

Differential scanning calorimetry measurements were performed on a DSC 1 of Mettler Toledo. All measurements are done in alumina pans at a heating rate of 10 K/min.

Single crystal X-ray diffraction was done with an Oxford Gemini diffractometer at 110 K applying Mo-K α (0.71 Å) and Cu-K α (1.54 Å) radiation. Structure resolution was done using a direct method by means of SHELXS-91.^[28] Structure refinement was performed with the software SHELX-97.^[29]

TG-MS measurements were recorded with a Thermo-Microscales TG 209 Iris[®] coupled with a mass spectrometer MS 403C Aëlos[®] from Netzsch. While heating up the sample a constant helium gasflow (30 ml·min⁻¹) is applied. As detector of the mass spectrometer a quadrupol-MS (QMA200) is used (measuring range $m/z = 0-300$).

Synthesis of adducts 6, 7 and 11

A solution of 15 mmol of imine base in 15 mL of dry diethyl ether was added drop-wise to an ice-cooled solution (–5 °C) of 30 mmol (45 mmol for **11**) of isocyanate in 40 mL dry diethyl ether. The solution was stirred for 12 h at room temperature. After removing the solvent by a rotational evaporator (25 °C), the crude product was gently recrystallized from a chloroform / *n*-hexane mixture and is obtained in yields ranging between 56 to 99 % (specific yields see Scheme 3). We also tested the auto polymerization behavior of MMA, MA and MAN under the used conditions. Results are provided in the Supporting Information in Table S5.

Synthesis of oligomers 13–15

A solution of 16.7 mmol diisocyanate in 15 mL dry diethyl ether was added drop-wise to an ice-cooled solution (–5 °C) of 16.7 mmol imine base in 15 mL of dry diethyl ether. After the addition of diisocyanate, the reaction mixture was allowed to

warm to room temperature and was stirred for another 18 h. After concentrating the solution to 10 mL by evaporating the solvent at reduced pressure, the reaction mixture was gently dripped into 400 mL of *n*-hexane under stirring of the *n*-hexane solution; a colorless precipitate of the imine base–diisocyanate oligomer was formed. After centrifugation of the suspension, the oligomer was washed for four times with 250 mL of *n*-hexane. The product is a foamy solid which is obtained in yields ranging between 37–88 % (details see Scheme 4) after drying under vacuum (5 mbar, room temperature, 12 h).

General acrylate polymerization procedure using adducts 6, 7 and 11 and oligomers 12–14

A solution of 200 mg adduct or oligomer dissolved in 5.3 mL (25 mmol) of MMA was stirred at 80 °C for 3 h. After polymerization, the viscous solution was cooled to room temperature, diluted in 30 mL of chloroform and slowly dripped into 300 mL of ice-cooled methanol under stirring. The precipitate was centrifuged off, dissolved in 15 mL of chloroform and re-precipitated in methanol. After repeating this process three more times, the polymer is dried in the vacuum oven (5 mbar) at room temperature for 12 h. Polyacrylate or polyacrylate-oligourea block-copolymer was isolated as colorless powder in yields ranging between 7 and 82 % (details see Table 5).

Products were characterized by ¹H / ¹³C NMR, FT-IR, DSC and elemental analysis, partially by TG-MS and ESI-MS. Measurement results are concluded in the Supporting Information section. Block co-polymerizations using MA were also performed after this method; the mixture was stirred under reflux at an oil bath temperature of 100 °C.

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Notes and references

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- 1 A. Wurtz, J. Liebig, *Ann. Chem.* 1849, **71**, 326.
- 2 O. Bayer, DRP 728981, DE patent, 1937.
- 3 E. Delebecq, J.-P. Pascault, B. Boutevin, F. Ganachaud, *Chem. Rev.* 2013, **113**, 80.
- 4 H. Ulrich, *Chemistry and Technology of Isocyanates*, John Wiley & Sons Ltd. Chichester, England 1996.
- 5 Z. Wicks, *Prog. in Org. Coatings* 1975, **3**, 73.
- 6 Z. Wicks, *Prog. in Org. Coatings* 1981, **9**, 3.
- 7 D. A. Wicks, Z. W. Wicks Jr., *Prog. in Org. Coatings* 1999, **36**, 148.
- 8 D. A. Wicks, Z. W. Wicks Jr., *Prog. in Org. Coatings* 2001, **41**, 1.

- 9 B. W. Kostyk, Z. W. Wicks Jr., *J. Polym. Sci.* 1979, **17**, 2423.
- 10 H. Kothandaraman, R. Thangavel, *J. Polym. Sci.: Part A: Polym. Chem.* 1993, **31**, 2653.
- 11 A. Mühlebach, *J. Polym. Sci.: Part A: Polym. Chem.* 1994, **32**, 753.
- 12 ^{a)} S. Subramani, Y. J. Park, Y. S. Lee, J. H. Kim, *Prog. in Org. Coatings* 2003, **48**, 71. ^{b)} X. Tassel, D. Barbry, L. Tighzert, *E. Polym. J.* 2000, **36**, 1745.
- 13 S. Subramani, I. W. Cheong, J. H. Kim, *Prog. in Org. Coatings* 2004, **51**, 329.
- 14 A. Sultan Nasar, S. Subramani, G. Radhakrishnan, *J. Polym. Sci.: Part A: Polym. Chem.* 1999, **37**, 1815.
- 15 ^{a)} H. J. Laas, R. Halpaap, J. Pedain, *J. Prakt. Chem.* 1994, **336**, 185. ^{b)} E. Querat, L. Tighzert, J. P. Pascault, K. Dušek, *Angew. Makromol. Chem.* 1996, **242**, 1.
- 16 S. Bode, M. Enke, H. Göris, S. Hoepfener, R. Weberskirch, M. D. Hager, U. S. Schubert, *Polym. Chem.* 2014, **5**, 2574.
- 17 H. Ulrich, B. Tucker, A. A. R. Sayigh, *Angew. Chem. Int. Ed.* 1968, **7**, 291.
- 18 R. Richter, *Tetrahedron Lett.* 1968, **48**, 5037.
- 19 R. Richter, *Chem. Ber.* 1968, **101**, 3002.
- 20 S. Ozaki, *Chem. Rev.* 1972, **72**, 457.
- 21 I. Polenz, F. G. Schmidt, S. Spange, *J. Polym. Sci.: Part A: Polym. Chem.* 2012, **50**, 1655.
- 22 I. Polenz, F. G. Schmidt, S. Spange, *J. Polym. Sci.: Part A: Polym. Chem.* 2013, **51**, 687.
- 23 J. Alsarraf, F. Robert, H. Cramail, Y. Landais, *Polym. Chem.* 2013, **4**, 904.
- 24 J. Alsarraf, Y. A. Ammar, F. Robert, E. Cloutet, H. Cramail, Y. Landais, *Macromolecules* 2012, **45**, 2249.
- 25 K. Schwetlick, *Organikum*, 23th Ed., Wiley-VCH Verlag, Weinheim, 2009.
- 26 J. Brandrup, E. H. Immergut, E. A. Grulke, A. Abe, D. R. Bloch, *Polymer Handbook*; Wiley-Interscience: New York, 1999.
- 27 M. Fineman, S. D. Ross, *J. Polym. Sci.* 1950, **2**, 259.
- 28 ^{a)} A. Altomare, G. Cascarano, C. Giacovazzo, A. Gualardi, *J. Appl. Cryst.* 1993, **26**, 343. ^{b)} G. M. Sheldrick, *Acta Cryst.* 2008, **A64**, 112.
- 29 ^{a)} G. M. Sheldrick, SHELXL-97, *Program for Crystal Structure Refinement*, University of Göttingen, 1997. ^{b)} G. M. Sheldrick, SHELXTL Version 5.1, *An Integrated System for Solving, Refining and Displaying Crystal Structures from Diffraction Data*, Siemens Analytical X-Ray Instruments, Madison, WI, 1990.