

Dalton Transactions

Accepted Manuscript



This is an *Accepted Manuscript*, which has been through the Royal Society of Chemistry peer review process and has been accepted for publication.

Accepted Manuscripts are published online shortly after acceptance, before technical editing, formatting and proof reading. Using this free service, authors can make their results available to the community, in citable form, before we publish the edited article. We will replace this *Accepted Manuscript* with the edited and formatted *Advance Article* as soon as it is available.

You can find more information about *Accepted Manuscripts* in the [Information for Authors](#).

Please note that technical editing may introduce minor changes to the text and/or graphics, which may alter content. The journal's standard [Terms & Conditions](#) and the [Ethical guidelines](#) still apply. In no event shall the Royal Society of Chemistry be held responsible for any errors or omissions in this *Accepted Manuscript* or any consequences arising from the use of any information it contains.

COMMUNICATION

Facile Synthesis of [(NHC)MCl(cod)] and [(NHC)MCl(CO)₂] (M= Rh, Ir) complexes

Cite this: DOI: 10.1039/x0xx00000x

R. Savka and H. Plenio*

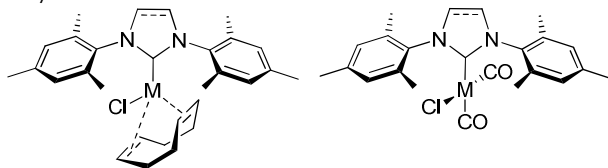
Received 00th September 2014,
Accepted 00th xxx 2014

DOI: 10.1039/x0xx00000x

www.rsc.org/

The reactions of [MCl(cod)]₂ (M= Rh, Ir) with different NHC·HX (X= Cl, I), K₂CO₃ in technical grade acetone under air provide simple access of various [(NHC)MX(cod)] complexes; a facile one-pot synthesis of [(NHC)MCl(CO)₂] (M= Rh, Ir) is also reported.

The use of *N*-heterocyclic carbenes has a profound impact on organometallic chemistry and homogeneous catalysis and during the last years metal complexes all across the periodic table have been synthesized.¹ [(NHC)RhCl(cod)] (cod = 1,5-cyclooctadiene) and [(NHC)RhCl(CO)₂] complexes were first reported by the Lappert group,² later Herrmann studied such complexes with Rh and Ir in some detail.³ Catalytic applications in transfer hydrogenation,⁴ hydrogenation,⁵ hydrosilylation⁶ or hydrogenation reactions for magnetization transfer⁷ are known. More important is the use of [(NHC)MCl(cod)] and [(NHC)MCl(CO)₂] (M= Rh, Ir) (scheme 1) for the determination of the electron donating properties (Tolman electronic parameter) of the respective NHC ligands,⁸ since Nolan et al.⁹ and Crabtree et al.¹⁰ have established the use of such complexes as practical alternatives to the traditional LNi(CO)₃ complexes.¹¹ Consequently, for a large proportion of newly prepared NHC ligands the respective [(NHC)MCl(cod)] and [(NHC)MCl(CO)₂] (M= Rh, Ir) have been synthesized.¹²

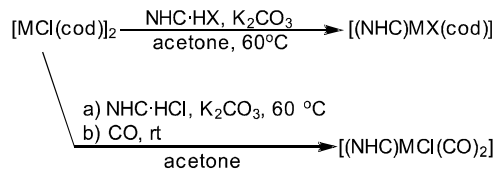


Scheme 1. [(NHC)MCl(cod)] and [(NHC)MCl(CO)₂] (M= Rh, Ir) complexes.

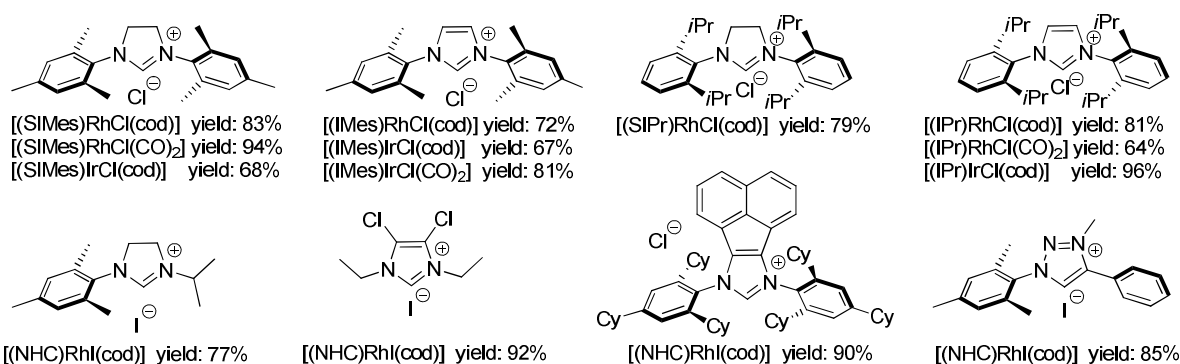
[(NHC)MX(cod)] (M= Rh, Ir) complexes (scheme 1) are prepared along two different routes: a) The direct reaction of the isolated or the in-situ generated carbene (obtained from treatment of the azolium salt

with a strong base) with [(MCl(cod))₂] appears to be the most popular approach.¹² Sometimes there is a different order in the addition of the reactants, such that [MCl(cod)]₂ is first treated with an alkoxide base (e.g. KOtBu, NaOMe) to generate the basic [(M(OR)(cod))₂] complex, followed by addition of the azolium salt.¹³ Whatever approach is chosen, anaerobic conditions are mandatory, furthermore base sensitive functional groups may cause problems. b) Alternatively, transmetalation reactions of [MCl(cod)]₂ with preformed Ag-NHC complexes^{10,14} or other NHC transfer reagents like imidazolium carboxylates,¹⁵ 2-(pentafluorophenyl)imidazolidines¹⁶ or imidazolidines¹⁷ have been employed. However, the generation of the respective Ag(NHC) salt or of the other NHC transfer reagents requires additional synthetic steps. The synthesis of the respective [(NHC)MCl(CO)₂] complexes relies on [(NHC)MCl(cod)], whose solutions are treated with CO gas, again this constitutes an additional reaction step.^{12f}

Simplicity in the synthesis of such complexes is essential for the application of the respective NHC-metal complexes. Therefore, we want to present here the facile one-step synthesis of [(NHC)MX(cod)] (X= Cl, I) and [(NHC)MCl(CO)₂] (M= Rh, Ir) under ambient atmosphere utilizing [MCl(cod)]₂ as the metal source. Our approach is motivated by recent publications from Gimeno et al.¹⁸ and Nolan et al.,¹⁹ who independently reported simple reaction conditions (acetone or CH₂Cl₂ and K₂CO₃ as the base) for the synthesis of various NHC-Au complexes from [AuCl(Me₂S)] or NHC-Cu complexes from CuX (X= Cl, Br, I).²⁰



Scheme 2. Simplified synthesis of [(NHC)MX(cod)] and [(NHC)MCl(CO)₂] (M= Rh, Ir; X= Cl, I).



Scheme 3. [(NHC)RhCl(cod)], [(NHC)RhI(cod)], [(NHC)RhCl(CO)₂], [(NHC)IrCl(cod)] and [(NHC)IrCl(CO)₂] complexes synthesized according to the improved synthesis protocol employing the depicted azolium salts.

The reactions of [RhCl(cod)₂] with SIMes-HCl using K₂CO₃ as the base in acetone was tested first and found to provide the desired [(SIMes)RhCl(cod)] in an isolated yield of 83% (scheme 2). In order to establish the generality of this approach seven additional azolium salts were tested (scheme 3). The nature of the NHC precursors was modulated (saturated vs. unsaturated; *N,N'*-diaryl vs. *N,N'*-aryl,alkyl vs. *N,N'*-dialkyl, backbone substituted and abnormal carbene) to cover a larger range of different NHC ligands. The respective eight different [(NHC)RhX(cod)] complexes were obtained in 72 – 92% yield. Notably, for all of the tested reactions, there is a larger than 90% conversion into [(NHC)RhX(cod)]. However, due to the small scale nature of the reactions (typically ca. 50 – 100 mg of product were formed), the isolated yields can be slightly lower, due to losses during work-up. The NHC transfer to Rh works equally well for azolium salt with an iodide counter ion and leads to the isolation of the respective [(NHC)RhI(cod)] complexes.

Next the reactions of [IrCl(cod)₂] and azolium salts leading to [(NHC)IrCl(cod)] were studied, by applying the same conditions as for the Rh complexes. For the unsaturated azolium salts (IMes-HCl, IPr-HCl) good yields of the respective [(NHC)IrCl(cod)] were obtained, employing the "Rh conditions". However, for the related saturated azolium salts (SIMes-HCl) additional optimization of the reaction conditions was required. The amount of base was found to be a critical parameter and the near stoichiometric use of base (1.2 equiv.) gives better results (scheme 3) than the 3 equivs. K₂CO₃ used for the synthesis of the Rh complexes.

Based on the high yielding conversion of [MCl(cod)₂] (M= Rh, Ir) into the respective [(NHC)MX(cod)], we became also interested, whether in-situ generated [(NHC)MCl(cod)] complexes can be reacted directly with CO to render the respective [(NHC)MCl(CO)₂] complexes (scheme 2). This one-pot approach turned out to be successful and several [(NHC)MCl(CO)₂] (scheme 3) were isolated in 64-94% yield. Due to the lower solubility of the respective carbonyl complexes, the isolated yields of those complexes can be higher than those of the related cod complexes.

In order to get some insight into the mechanism of the facile [(NHC)MX(cod)] formation, SIMes-HCl was treated with K₂CO₃ in acetone. In the absence of metal complexes, this reaction slowly led to the respective ring-opened formamides, which are known decomposition products of the free NHC ligands in the presence of

moisture.²¹ Nonetheless, we consider it unlikely that the [(NHC)MX(cod)] complexes are formed via free carbenes, since free carbenes are unstable in acetone. Therefore the reaction of [RhCl(cod)₂] with K₂CO₃ in acetone was studied, which leads to the rapid formation of a yellow precipitate. This solid was isolated and reacted with SIMes-HCl. Notably, in the absence of additional base, the formation of [(SIMes)RhCl(cod)] was observed. This precipitate appears to be the key species in the synthesis and it has to contain a basic group. Based on literature data the yellow solid can either be the carbonato complex²² or the hydroxo complex [Rh(OH)(cod)]₂.²³ The latter complex was reported to form in 95% yield from [RhCl(cod)₂] in water/KOH.²³ Neither the IR spectrum of the yellow material nor its ¹³C NMR provide evidence for the formation of the carbonato complex. Instead, the ¹H NMR and the IR spectrum display characteristic signals for the well known [Rh(OH)(cod)]₂.²³ Based on this we assume, that the primary product of the reaction is the hydroxo complex, which directly reacts with an azolium salt to form the respective NHC-M (M= Rh, Ir) complexes. The better stability of such complexes with Rh than with Ir²³ may explain the slightly better synthetic results obtained for Rh complexes.

In conclusion, we have demonstrated the facile one-step synthesis of [(NHC)MX(cod)] (M= Rh, Ir; X= Cl, I) and the simple one-pot synthesis of [(NHC)MCl(CO)₂] complexes for different azolium salts. The desired complexes are produced in excellent yields with different NHC ligands. The reaction of the metal precursor [MCl(cod)₂] with the respective azolium salt is carried out in technical grade acetone under ambient atmosphere and provides easy access of the respective NHC complexes – for the synthesis of the respective carbonyl complexes, CO is bubbled through the crude reaction mixture.

The authors gratefully acknowledge support of this work by the DFG via DFG 178/13-2. We wish to thank one referee for useful literature references.

Notes and references

^a Organometallic Chemistry, Alarich-Weiss-Str. 12, Fachbereich Chemie, TU Darmstadt, 64287 Darmstadt, Germany, e-mail: plenio@tu-darmstadt.de

† General procedure for the synthesis of [(NHC)RhX(cod)] (X = Cl, I), and [(NHC)IrCl(cod)] complexes: A vial was charged, under air, with the corresponding NHC·HX (1 equiv), [MCl(cod)]₂ (0.5 equiv) and K₂CO₃ (3 equiv). The resulting mixture was suspended in acetone (3 mL) and stirred for 20 h at 60 °C (for one-pot synthesis of carbonyl complexes, continue below →). After this time the volatiles were removed *in vacuo* and CH₂Cl₂ was added (3 mL). The mixture was filtered through a pad of silica. The pad of silica was washed with CH₂Cl₂ until the filtrate becomes colorless and the product dried *in vacuo*.

→ One-pot synthesis of [(NHC)MCl(CO)₂] (M = Rh, Ir): Next, the mixture was cooled to room temperature and carbon monoxide bubbled through under vigorous stirring for 20 min. The volatiles were removed *in vacuo*, the residue dissolved in CH₂Cl₂ and filtered through Celite, which was washed with CH₂Cl₂ until the filtrate becomes colorless. The volatiles were removed *in vacuo* and the residue washed with pentane (5 mL). The pale yellow precipitate was collected by filtration, washed with pentane and dried *in vacuo*.

Electronic Supplementary Information (ESI) available: Experimental details and full characterisation of complexes. See DOI: 10.1039/c000000x/

- 1 *N-Heterocyclic Carbenes*, (ed. S. P. Nolan), Wiley-VCH, Weinheim 2014.
- 2 a) M. J. Doyle and M. F. Lappert, *Chem. Commun.*, 1974, 679-680; b) J. E. Hill and T. A. Nile, *J. Organomet. Chem.*, 1977, **137**, 293-300; c) M. J. Doyle, M. F. Lappert, P. L. Pye and P. Terreros, *J. Chem. Soc., Dalton Trans.*, 1984, 2355-2364.
- 3 a) W. A. Herrmann, M. Elison, J. Fischer, C. Köcher and G. R. J. Artus, *Chem. Eur. J.*, 1996, **2**, 772-780; b) W. A. Herrmann, J. Schütz, G. D. Frey and E. Herdtweck, *Organometallics*, 2006, **25**, 2437-2448.
- 4 S. Gülcemal, A. G. Gökçe and B. Çetinkaya, *Inorg. Chem.*, 2013, **52**, 10601-10609.
- 5 E. L. Kolychev, S. Kronig, K. Brandhorst, M. Freytag, P. G. Jones and M. Tamm, *J. Am. Chem. Soc.*, 2013, **135**, 12448-12459.
- 6 A. Monney and M. Albrecht, *Chem. Commun.*, 2012, **48**, 10960-10962.
- 7 M. J. Cowley, R. W. Adams, K. D. Atkinson, M. C. R. Cockett, S. B. Duckett, G. G. R. Green, J. A. B. Lohman, R. Kerssebaum, D. Kilgour and R. E. Mewis, *J. Am. Chem. Soc.*, 2011, **133**, 6134-6137.
- 8 a) S. Wolf and H. Plenio, *J. Organomet. Chem.*, 2009, **694**, 1487-1492; b) T. Dröge and F. Glorius, *Angew. Chem. Int. Ed.*, 2010, **49**, 6940-6952; c) D. J. Nelson and S. P. Nolan, *Chem. Soc. Rev.*, 2013, **42**, 6723-6753.
- 9 N. Boens, V. Leen and W. Dehaen, *Chem. Soc. Rev.*, 2011, **41**, 1130-1172.
- 10 A. R. Chianese, X. Li, M. C. Janzen, J. W. Faller and R. H. Crabtree, *Organometallics*, 2003, **22**, 1663-1667.
- 11 a) R. Dorta, E. D. Stevens, N. M. Scott, C. Costabile, L. Cavallo, C. D. Hoff and S. P. Nolan, *J. Am. Chem. Soc.*, 2005, **127**, 2485-2495; b) C. A. Tolman, *Chem. Rev.*, 1977, **77**, 313-348.
- 12 a) C. D. Varnado, E. L. Rosen, M. S. Collins, V. M. Lynch and C. W. Bielawski, *Dalton Trans.*, 2013; b) A. Makhloufi, W. Frank and C. Ganter, *Organometallics*, 2012, **31**, 2001-2008; c) H. Valdés, M. Poyatos and E. Peris, *Organometallics*, 2013, **32**, 6445-6451; d) V. César, N. Lugan and G. Lavigne, *Eur. J. Inorg. Chem.*, 2010, **2010**, 361-365; e) S. Dierick, D. F. Dewez and I. E. Markó, *Organometallics*, 2014, **33**, 677-683; f) S. Leuthäuser, D. Schwarz and H. Plenio, *Chem. Eur. J.*, 2007, **13**, 7195-7203; g) Y. Borguet, G. Zaragoza, A. Demonceau and L. Delaude, *Dalton Trans.*, 2013, **42**, 7287-7296; h) A. Fürstner, M. Alcarazo, H. Krause and C. W. Lehmann, *J. Am. Chem. Soc.*, 2007, **129**, 12676-12677; i) M. Mayr, K. Wurst, K.-H. Ongania and M. R. Buchmeiser, *Chem. Eur. J.*, 2004, **10**, 1256-1266; j) M. Iglesias, D. J. Beetstra, A. Stasch, P. N. Horton, M. B. Hursthouse, S. J. Coles, K. J. Cavell, A. Dervisi and I. A. Fallis, *Organometallics*, 2007, **26**, 4800-4809; k) R. D. Savka and H. Plenio, *J. Organomet. Chem.*, 2012, **710**, 68.; l) with anhydrous K₂CO₃: T. Sato, Y. Hirose, D. Yoshioka, S. Oi, *Organometallics* 2012, **31**, 6995-7003.
- 13 a) M. T. Zamora, M. J. Ferguson, R. McDonald, M. Cowie, *Organometallics* 2012, **31**, 5463-5477; b) O. Torres, M. Martín, E. Sola, *Organometallics* 2009, **28**, 863-870; c) I. Özdemir, B. Yigit, B. Çetinkaya, D. Ülkü, M. N. Tahir, C. Arıcı, *J. Organomet. Chem.* 2001, **633**, 27-32.
- 14 a) A. R. Chianese, A. Kovacevic, B. M. Zeglis, J. W. Faller and R. H. Crabtree, *Organometallics*, 2004, **23**, 2461-2468; b) S. Saravanakumar, A. I. Oprea, M. Kindermann, P. G. Jones and J. Heinicke, *Chem. Eur. J.*, 2006, **12**, 3143-3154.
- 15 A. M. Voutchkova, M. Feliz, E. Clot, O. Eisenstein and R. H. Crabtree, *J. Am. Chem. Soc.*, 2007, **129**, 12834-12846.
- 16 A. P. Blum, T. Ritter and R. H. Grubbs, *Organometallics*, 2007, **26**, 2122-2124.
- 17 A. Prades, M. Poyatos, J. A. Mata and E. Peris, *Angew. Chem. Int. Ed.*, 2011, **50**, 7666-7669.
- 18 R. Visbal, A. Laguna and M. C. Gimeno, *Chem. Commun.*, 2013, **49**, 5642-5644.
- 19 A. Collado, A. Gomez-Suarez, A. R. Martin, A. M. Slawin and S. P. Nolan, *Chem. Commun.*, 2013, **49**, 5541-5543.
- 20 O. Santoro, A. Collado, A. M. Z. Slawin, S. P. Nolan and C. S. J. Cazin, *Chem. Commun.*, 2013, **49**, 10483-10485.
- 21 M. Emin Günay, N. Özdemir, M. Ulusoy, M. Uçak, M. Dinçer and B. Çetinkaya, *J. Organomet. Chem.*, 2009, **694**, 2179-2184.
- 22 E. G. Lundquist, K. Folting, J. C. Huffman and K. G. Caulton, *Inorg. Chem.*, 1987, **26**, 205-208.
- 23 R. Uson, L. A. Oro, J. A. Cabeza, H. E. Bryndza and M. P. Stepro, *Inorg. Synth.*, 1985, **23**, 126-130.

