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Calcium Carbonate Crystal Growth Beneath Langmuir Monolayers of Acidic β-Hairpin Peptides

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Abstract

Four amphiphilic peptides with designed hairpin structure were synthesized and their monolayers were employed as model systems to study biologically inspired calcium carbonate crystallization. Langmuir monolayers of hairpin peptides were investigated by surface pressure area isotherms, surface potential isotherms, Brewster angle microscopy (BAM), atomic force microscopy (AFM) and Fourier transform infrared (FTIR) spectroscopy. A β-hairpin conformation was found for all peptides at the air/water interface although their packing arrangements seem to be different. Crystallization of calcium carbonate under these peptide monolayers was investigated at different surface pressures and growth times both by *in situ* optical microscopy, BAM and *ex situ* investigations such as scanning electron microscopy (SEM) and transmission electron microscopy (TEM). An amorphous calcium carbonate precursor was found at the initial crystallization stage. The crystallization process occurred in three stages. It starts from the nucleation of amorphous particles being a kinetically controlled process. Crystal nuclei subsequently aggregate to large particles and vaterite crystals start to form inside the amorphous layer, with the monolayer fluidity exerting an important role. The third process includes the re-crystallization of vaterite to calcite, which is thermodynamically controlled by monolayer structural factors including the monolayer flexibility and packing arrangement of the polar headgroups. Thus, the kinetic factors, monolayer fluidity and flexibility as well as structure factors govern the crystal morphology and polymorph distribution simultaneously and synergistically.

Introduction

Living organisms are capable to build intricate shells and casings from calcium carbonate which are widely known as biominerals.[1, 2] Most of these biominerals are structurally well-organized composites of inorganic and organic components. Among the most frequently studied biominerals, nacre, the lustreous inner part of many mollusc shells, consists of alternating layers of aragonite crystal tablets that are interspersed with thin sheets of an organic matrix, the latter consisting of a complex mixture of proteins.[3] These protein layers inside the mineral phase not only act as a structure component but also have been recognized to be a template or modifier of calcium carbonate mineralization.[4, 5] Such an unusual "brick-and-mortar" structure makes nacreous biominerals exceptionally tough and resistant.[6] Some proteins isolated from the nacreous parts of mollusc biominerals are particularly rich in acidic amino acid residues, e.g. aspartic or glutamic acid.[7] *In vitro* experiments provide strong evidence for the hypothesis that these acidic peptides are directly involved in the mineralization process and play a crucial role in modulating the mineral crystallization process and thus determine the final morphology or polymorph of the minerals.[3, , 9, 11 12, 15]. Moreover, it was proposed that the uniform growth direction of calcium carbonate crystals induced by the biological organic matrix might relate to structurally well-organized packing motifs of carboxylate residues emanating from β-sheet domains of the acidic proteins.[16] However, a direct structural relationship, in the sense of an expitaxial match between these self-organizing peptides and the juxtaposed $CaCO_3$ crystal face has been subject of a controversial debate.[17]

Interfacial properties of polypeptides in living organisms seem to play a key role in modifying calcium carbonate crystal growth since it is assumed that the initial crystal nucleation takes place at peptide-modified surfaces or interfaces. However, only few investigations employed artificial peptides as $CaCO_3$ crystal growth modifiers, which included water-soluble peptides,[18, 19, 20]and water-insoluble Langmuir monolayers,[21, 22] Insoluble monolayers of peptides in particular might form an organized supramolecular template matrix with a well-defined two-dimensional structure, thus providing a potential means to control the nucleation and crystallization of inorganic crystals. From an experimentalist's point of view it is straight-forward to adjust monolayer properties such as molecular conformation and/or charge density by changing the experimental conditions (i.e. monolayer surface pressure, pH and ion conc. of the aqueous subphase). The crystallization beneath peptide monolayers can be initiated at different conditions and in a controlled manner. In addition, the two-dimensional monolayer matrix structure can be characterized structurally, both by *in situ* methods and by *ex situ* characterization of the transferred Langmuir-Blodgett films, finally revealing the key factors which determine the surface-initiated crystallization process. Langmuir monolayers of peptides thus constitute reductionistic model systems for peptide-induced biomineralisation, i.e. to investigate crystal growth at self-organized organic matrices.

However, although monolayer-controlled crystal growth might seems to be a simplistic

model system, diverse and partly contradictory hypotheses can be found in literature about calcium carbonate crystallization beneath monolayers of amphiphilic molecules covering a wide range of different molecular structures. In seminal studies, Mann and Heywood *et al.* attributed the oriented nucleation to the geometrical and/or stereochemical match between the monolayer packing arrangement and the oriented mineral surface.[23, 26] Later, a multistep assembly process from an amorphous to a crystalline phase was suggested by Xu *et al.*[27] Other groups demonstrated that the CO_2 evaporation rate effectively influences crystal growth beneath monolayers, which may add to or even overlay the template effect of the monolayer.[28] *In situ* synchrotron X-ray diffraction studies have demonstrated that additives such as poly(sodium acrylate) and Mg^{2+} ions dissolved in the aqueous subphase lead to changes in the kinetics of the crystallization process.[29] Employing a vast range of structurally different, tailored amphiphiles it was concluded by Volkmer *et al.* that non-specific factors such as the monolayer's surface charge density or its effective dipole moment seems to be key factors that determine the morphology and polymorph of the crystals grown underneath the monolayer.[30, 33] In addition, monolayer flexibility and its corresponding dynamic structural changes during crystallization were shown to be crucial factors in the mineralization process.[34]

Continuing these model studies we here report on $CaCO_3$ crystallization underneath monolayers of amphiphilic, structurally well-defined β-hairpin peptides which were chosen for reasons of structural analogy with natural polyacidic peptides often found in biominerals. These hairpin peptides are composed of alternating hydrophobic and hydrophilic amino acid residues that spontaneously fold into a D-Pro-Gly-induced β-hairpin conformation at the air/water interface not only by interfacial assembly but also by inter-strand hydrogen bonding interactions.[35-37] Artificial amphiphilic hairpin peptides represent a novel class of functional model peptides that form insoluble monolayers. The film-forming properties of these molecules as yet have not been studied in depth with respect to different in-film molecular packing arrangements. The possibility to adjust the monolayer's surface charge density in conjunction with different molecular packing arrangements provides additional degrees of experimental freedom if compared to single strand model peptides comprised of alternating hydrophobic / hydrophilic residues. From an experimental point of view, the latter peptides, although structurally more closely related to natural proteins, are very difficult to prepare by standard solid-phase synthesis owing to their high propensity to form insoluble aggregates, which furthermore renders them difficult to spread as monolayer films in a controlled and reproducible fashion. Calcium carbonate crystallization beneath β-hairpin peptide monolayers having different amino acid sequences will be described in the following. Although these peptides have similar primary structure, it is found that $CaCO_3$ nucleation and crystallization beneath each monolayer they from follow distinct pathways, which might be ascribed to the different peptide packing arrangements at the air/water interface.

Experimental Section

General All chemicals were purchased from *Sigma Aldrich* (Hamburg, Germany), *Acros* (Geel, Belgium), *Alfa Aesar* (Ward Hill, USA) and *VWR* (Darmstadt, Germany) and were employed without additional purification. Dimethylformamide was distilled in vacuum from ninhydrin. 2-Chlorotrityl resin (0.6 mmol/g), coupling reagents and all Fmoc-amino acids were purchased from *Iris Biotech* (Marktredwitz, Germany) and *Orpegen* (Heidelberg, Germany). Peptide synthesis was performed manually. Preparative RP-HPLC was done on a Thermo Separation Products system equipped with a Vydac 218 TP 1022 efficiency column (22 × 250 µm, 10 µm) with water/acetonitrile gradients as the eluent and UV detection at 220 nm. Analytical RP-HPLC was performed on a Thermo Separation Products system equipped with a Phenomenex Jupiter 4 µ Proteo 90 Å column (250 x 4.6 mm, 4 µm) with water/acetonitrile gradients as the eluent and UV detection at 220 nm.

MALDI-ToF mass spectra were measured on a *Voyager DE Instrument* (*PE Biosystems*, Weiterstadt, Germany) mounted with a 1.2 m flight tube. 2,5-Dihydroxybenzoic acid was used as the matrix. Depending on the mass range the ions were accelerated at 15 to 25 kV with the option of detecting positive or negative ions. The instrument default calibration was used for calibrating the mass axis.

Manual Solid Phase Peptide Synthesis – General Procedure

Resin loading

The *o*-chlorotrityl resin was loaded in dry CH_2Cl_2 (10 mL per g Fmoc-amino acid) in the presence of DIPEA (3 equiv) with Fmoc-Asp(OtBu)-OH (3 equiv; for peptides **1a** and **2a**) or with Fmoc-Glu(OtBu)-OH (3 equiv) for peptides **1b** and **2b**. After shaking the mixture at room temperature for 5 min, additional 4.5 equiv DIPEA were added. The resin was shaken gently at room temperature for another 60 min, before methanol (0.8 mL methanol per g resin) was added to the reaction mixture in order to solvolyze unreacted *o*-chlorotrityl chloride on the resin. The resin loaded with Fmoc-amino acid was then filtered and subsequently washed with CH_2Cl_2 (4×), DMF (4×), CH_2Cl_2 (4×), and CH_3OH (4×). After drying in vacuum overnight the resin loading was determined UV-spectroscopically by cleaving Fmoc from a small resin sample with piperidine and measuring the absorbance of the piperidine-dibenzofulvene adduct at 290 nm.

Peptide Coupling

The resin loaded with the first amino acid (resin loading 800 µmol amino acid/g resin, 250-300 µmol) or the peptidyl resin in the course of the synthesis was treated with piperidine/DBU/DMF (2:2:96, v/v) for 10 min at room temperature to cleave the Fmoc group. The resin was washed with CH_2Cl_2 (4×), DMF (4×), CH_2Cl_2 (4×), CH_3OH (4×) and dried under vacuum overnight.

For peptide elongation the corresponding Fmoc-amino acid (3-4 equiv with respect to resin loading)

was dissolved in a minimum volume of DMF (1 mmol amino acid/mL) and pre-activated for up to 5 min with 3-4 equiv of TBTU, dissolved in a minimum volume of DMF, and 6 equiv DIPEA were added to the reaction mixture. The pre-activated Fmoc-amino acid was then added under argon to the resin loaded with the first amino acid or peptide and agitated for 45 minutes at room temperature. The reaction progress result was monitored by MALDI-ToF MS or Kaiser test.

For capping of unreacted amino groups 3 equiv Ac_2O (based on resin loading) and 3 equiv Et_3N, dissolved in DMF, were added and the mixture was agitated at room temperature for 30 min. Then the resin was washed with CH_2Cl_2 (4×), DMF (4×), CH_2Cl_2 (4×).

After assembly of the protected peptide on the resin, it was cleaved by adding 5 mL TFA in CH_2Cl_2 (1%, v/v). After 5 min the resin was filtered and the filtrate was neutralized with pyridine. The cleavage was repeated once and the combined filtrates were evaporated to dryness under reduced pressure. The crude fully protected peptides were lyophilized.

Deprotection of the protected peptides was achieved by treatment with the cleavage cocktail TFA/H_2O/TIS (95:2.5:2.5) for several hours at ambient temperature. The crude products were purified by preparative HPLC using acetonitrile/water/TFA gradients. The pure peptides were obtained in low yields because of their high tendency towards aggregation and precipitation.

H-Asp-Phe-Asp-Phe-Asp-Phe-Asp-D-Pro-Gly-Asp-Phe-Asp-Phe-Asp-Phe-Asp-OH (**1a**)

Yield: 20.5 mg, 1.0 µmol (0.33%); $C_{93}H_{106}N_{16}O_{33}$; 1977.94 g/mol

MS (MALDI-ToF): m/z = 1978.37 [M+H]+, 1999.42 $[M+Na]^+$, 2015.59 $[M+K]^+$

Calculated Mass (monoisotopic): $[M+H]^+$ = 1977.73, $[M+Na]^+$ = 1999.71, $[M+K]^+$ = 2015.69

H-Glu-Phe-Glu-Phe-Glu-Phe-Glu-D-Pro-Gly-Glu-Phe-Glu-Phe-Glu-Phe-Glu-OH (**1b**)

Yield: 8.5 mg (4.1 µmol; 1.6%); $C_{101}H_{122}N_{16}O_{33}$; 2088.14 g/mol

MS (MALDI-ToF): m/z = 2086.92 $[M+H]^+$, 2108.32 $[M+Na]^+$, 2124.35 $[M+K]^+$,

Calculated Mass (monoisotopic): $[M+H]^+$ = 2087.84, $[M+Na]^+$ = 2109.83, $[M+K]^+$ = 2125,80

H-Asp-Phe-Ser-Phe-Asp-Phe-Ser-D-Pro-Gly-Asp-Phe-Ser-Phe-Asp-Phe-Ser-OH (**2a**)

Yield: 15.5 mg (8.3 µmol; 3.3%); $C_{89}H_{106}N_{16}O_{29}$; 1863.88 g/mol

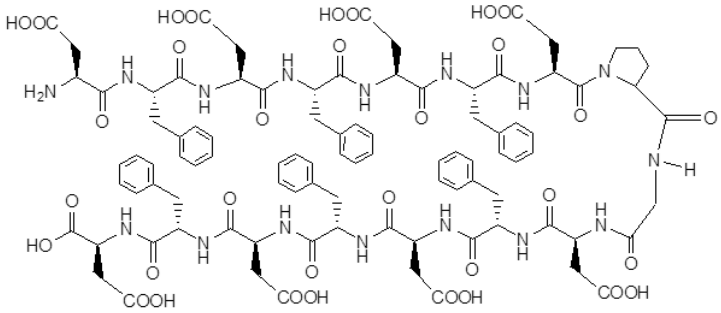
MS (MALDI-ToF): m/z = 1886.83 $[M+Na]^+$; calculated mass (monoisotopic): $[M+Na]^+$ = 1885.72

H-Glu-Phe-Ser-Phe-Glu-Phe-Ser-D-Pro-Gly-Glu-Phe-Ser-Phe-Glu-Phe-Ser-OH (**2b**)

Yield: 11.2 mg (5.8 µmol; 2.3%); $C_{93}H_{114}N_{16}O_{29}$; 1918.79 g/mol

MS (MALDI-ToF): m/z = 1941.36 $[M+Na]^+$, 1957.61 $[M+K]^+$; calculated mass (monoisotopic): $[M+Na]^+$ = 1941.78, $[M+K]^+$ = 1957.76

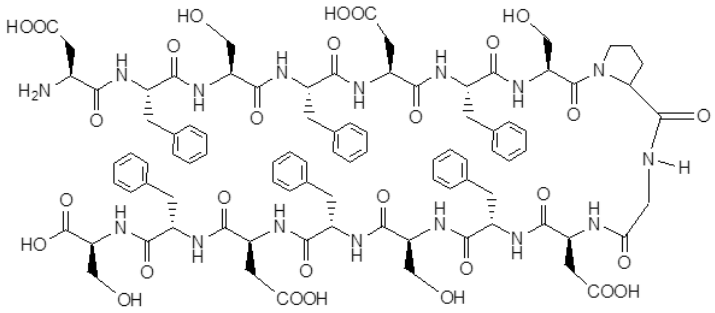
CD spectroscopy CD spectra were obtained on a *J-810* spectrometer equipped with a *CDF-4265 Peltier* unit for temperature control (*Jasco*, Groß-Umstadt, Germany). The spectra were recorded using a 0.2 mm path length quartz cell in the range of 185-300 nm at a scanning rate of 50 nm min^{-1} with three accumulations, a data pitch of 0.2 nm, a spectral band width of 1 nm and a response time of 1 s. Molar ellipticity per amino acid residue $[\theta]_{mrw}$ was calculated as follows: $[\theta]_{mr} = \theta/(10 \cdot N \cdot c \cdot l)$. θ represents the ellipticity in millidegrees, N the number of amino acid residues, c the molar concentration in mol L^{-1} and l the cell path length in cm. The CD spectra were smoothened using the means movement algorithm.



1a

H-Xaa-(Phe-Xaa)$_3$-D-Pro-Gly-Xaa-(Phe-Xaa)$_3$-OH

1a: *Xaa = Asp* **1b:** *Xaa = Glu*



2a

H-Xaa-Ph-Ser-Phe-Xaa-Phe-Ser-D-Pro-Gly-Xaa-Phe-Ser-Phe-Xaa-Phe-Ser-OH

2a: *Xaa = Asp* **2b:** *Xaa = Glu*

Scheme 1. Molecular structures and primary sequences of the hairpin peptides.

Langmuir monolayer investigations

Measurements of surface pressure area (π-A) isotherms and film depositions were carried out on a double barrier. Langmuir balance (NIMA 702 BAM, Nima Technology, UK). The Teflon trough has large dimensions (72 cm long × 10 cm wide × 0.5 cm deep). The volume of the subphase is about 360 mL in all experiments.

The surface pressure of the monolayers was measured using a filter paper Wilhelmy plate. The Langmuir film was formed by spreading a chloroform solution of peptides (0.2 mg/mL, with 10% DMF) onto Milli-Q purified water (ρ >18 MΩ•cm). After evaporation of the solvent (approx. 15 min), π-A isotherms of the spread films were recorded at room temperature (approx. 22°C) by compressing or expanding the Langmuir film at a barrier speed of 15 cm^2/min. Surface potential measurements were performed with a Trek electrometer. A vibrating electrode was placed about 2 mm above the air/water interface and a stainless steel plate as the reference electrode was immersed in the subphase. The potential solubilisation of peptides in the aqueus subphase was tested by measuring at least 3 repeated compression/decompression cycles of all monolayer samples (pure water or 9 mM $CaCl_2$). In no case any shift of the limiting area up to a maximum compression of 20 mN/m was found.

Monolayers and multilayers of peptides were transferred vertically at 15 mN/m on freshly cleaved mica, CaF_2, and silicon substrates for AFM, FT-IR and XPS measurements, respectively. The transfer speed was 2 mm/min for both upwards and downwards strokes.

Crystal growth experiments

Solutions of calcium bicarbonate were prepared by bubbling carbon dioxide gas through a stirred aqueous solution of $Ca(HCO_3)_2$ (9 mM) for a period of 2h. The monolayer was formed by adding 0.2 mg/mL of peptide solution (chloroform/DMF, 90%/10% vol.) drop-wise onto the surface of freshly prepared calcium bicarbonate solution followed by compression to a pressure of 0.1mN/m. Crystal growth underneath the monolayer was studied at several time intervals either by *in situ* optical microscopy, Brewster angle microscopy or by scanning electron microscopy (SEM) and transmission electron microscopy (TEM). For SEM the crystals floating at the air/water interface were transferred onto glass cover slips laid on the monolayer. TEM samples were prepared by carefully contacting the Cu grids with the monolayer and subsequent lift-off.

Analytical Characterization Techniques

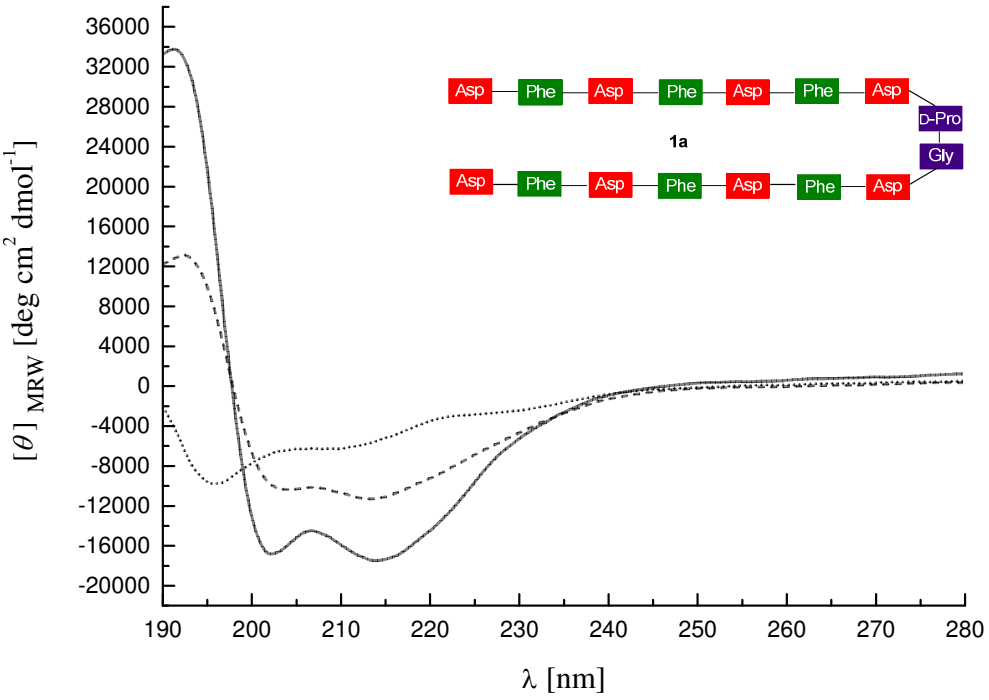
Infrared spectra were obtained using a Bruker IFS113v spectrometer. Raman spectroscopy was performed with a Horiba Jobin Yvon spectrometer. Brewster angle microscopy (BAM) was carried out with a NIMA Langmuir trough (NIMA 702 BAM) using a BAM-2 (I-Elli2000 supplied by NFT, Nanofilm Technologie, Göttingen, Germany). Images of the films were recorded with a lateral resolution of 2 μm. Crystal morphologies were transferred onto cover slides for SEM observation. Scanning electron microscopy (SEM) was carried out on Hitachi Tabletop 1000. Atomic force microscopy (AFM) was performed with the Agilent 5500 AFM and standard scanning probes from Nanosensors (PPP-FMR) with radii typically smaller than 7nm (nominal force constant k=2.8N/m).

Transmission electron microscope (TEM) investigations were carried out on a Philips CM20 operating at 200 kV. X-ray photoelectron spectra (XPS) were acquired with a Physical Electronics PHI 5800 Multi ESCA system with monochromic Al K_α radiation.

Results and Discussion

Conformational Analysis of the Hairpin Peptides by CD Spectroscopy

The solution conformation of peptides **1a** and **1b**, respectively was investigated in hexafluoroisopropanol (HFIP) and HFIP/H_2O solutions. Halogenated alcohols such as TFE and HFIP are widely applied solvents for peptides in CD spectroscopy because of their excellent solubilizing properties for some otherwise hardly soluble peptides such as β-sheet derivatives, and their inclination to stabilize ordered structures, e.g. α-helical conformations in peptides. Structures containing turns, β-hairpins, β-sheets, and hydrophobic clusters are also observable in the presence of TFE or HFIP. HFIP is a better hydrogen bond donor but a poorer hydrogen bond acceptor compared to water.[38] HFIP preferentially serves as a hydrogen bond donor interacting with the main chains of the peptide. This solvation preference of HFIP leads to enhanced intra-chain amide to carbonyl hydrogen bonding.



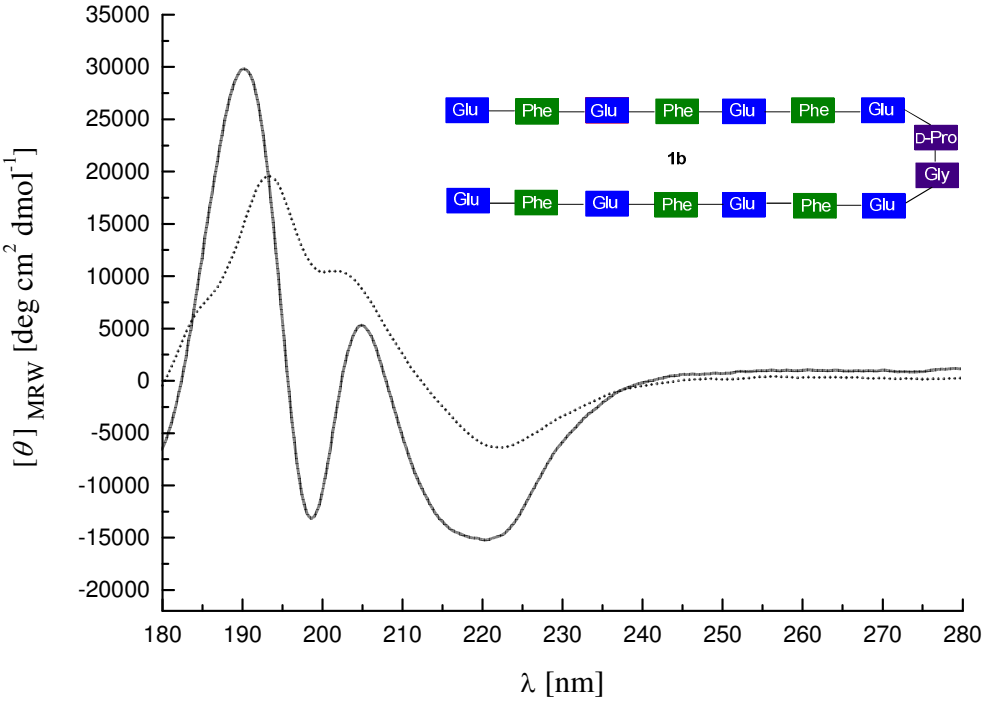


Figure 1. a) CD spectra of **1a** (0.5 mg/mL in HFIP, dotted line; 0.5 mg/mL in HFIP/H_2O 1:1, dashed line; 0.33 mg/mL in HFIP/H_2O 1:2, solid line), b) CD spectra of **1b** (0.3 mg/mL in HFIP, dotted line; 0.15 mg/ml in HFIP/H_2O 1:1, solid line)

The peptides were designed as β-hairpins and their folded conformation is based on three stabilizing factors: i) Hydrogen bonds between backbone amide moieties across the strands, ii) side chain interactions across two β-strands, and iii) the presence of a reverse turn (β-turn). In particular, the interaction between aromatic side chains forming hydrophobic clusters as a stabilizing effect has been shown for a number of β-hairpin peptides.[39]

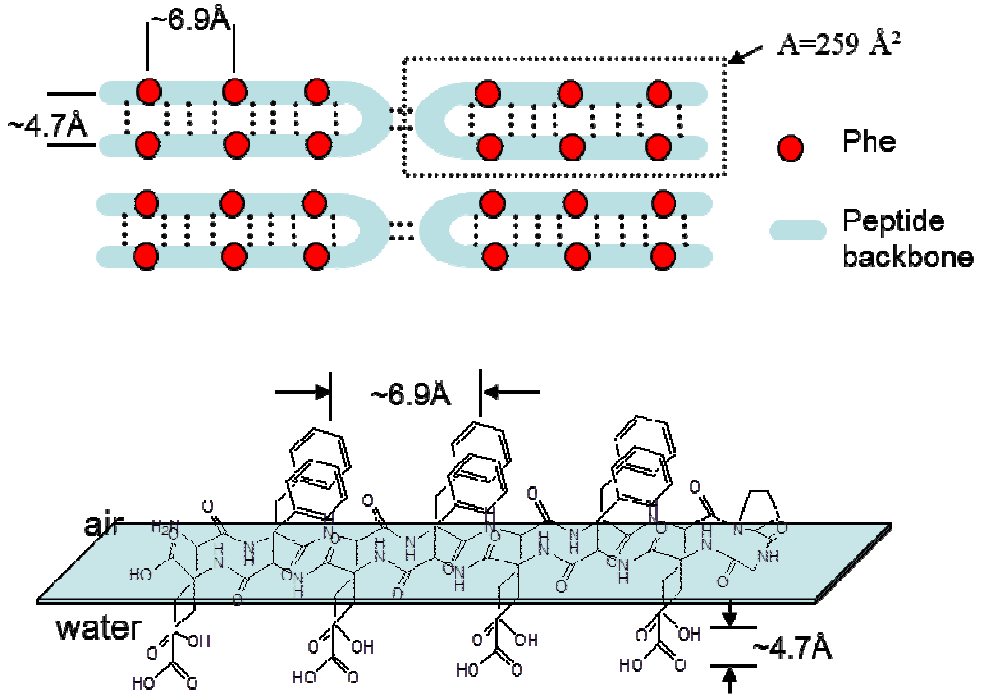
The CD spectrum of **1a** in HFIP indicates the predominance of random coil conformation, with a negative band at 195 nm (see Figure 1.a), corresponding to an unordered structure. Solvents like TFE in principle promote sheet or helix formation by reducing the water activity.[40] However, the overall influence depends on the structure propensity of the sequence. Fluorinated alcohols like TFE or HFIP have been found to attenuate hydrophobic interactions between residues that are distant in the peptide sequence, like e.g. across strands in some β-sheets, and, therefore, are able to disrupt these structures.[38] In contrast to peptide **1a**, the CD spectrum of **1b** in HFIP indicates the predominance of a β-hairpin, composed of β-sheet and β-turn conformations, with the typical positive band at 194 nm and the negative band at 221 nm, as shown in Figure 1b. The positive shoulder at 202 nm that increased with concentration (data not shown) is characteristic of the β-turn present in the β-hairpin conformation.

A conformational change was observed for both peptides upon addition of water to the HFIP solutions. The CD spectrum of the Asp-containing peptide **1a** displayed increasing β-hairpin

characteristics (positive band at 194 nm, increase in negative intensity at 215 nm) with increasing water contents (Figure 1a). Likewise, the CD-spectrum of the Glu-containing peptide **1b** showed a positive band at 190 nm and a negative band at 220 nm, also indicating β-hairpin conformation (Figure 1b). However, for both peptides **1a** and **1b** the positive CD effects at 190-194 nm are accompanied by negative bands at 200-203 nm and additional positive bands at 207-208 nm. This unusual pattern in aqueous solutions can be assigned to an interaction of aromatic chromophores. Such effects on hairpin stability have in particular been described for specific Trp/Trp, Trp/Tyr, and Tyr/Tyr interactions.[39] It has been suggested that they efficiently stabilize the secondary structure and can be employed to monitor the cross-strand interaction between the aromatic side chains. Moreover, such orthogonal arrangements of aromatic rings have been observed in proteins and have been suggested as a source of tertiary structure stabilization.[41] This shows that the increase of the water concentration stabilizes the β-hairpin conformation especially of peptide **1b**, presumably by hydrophobic interactions between the phenylalanine side chains.

Monolayer Properties of Hairpin Peptides

Figure 2 shows the π-A isotherms of all peptides on water and $CaCl_2$ subphase at room temperature (*ca.* 22°C). The isotherms of the peptides **1a** and **1b** on the water subphase show two liquid expanded stages separated by a phase transition. The limiting area of the peptides on the water subphase is 225 $Å^2$ for **1a,** and 250 $Å^2$ for **1b**, respectively.



Scheme 2. β-Hairpin packing arrangement of the peptides at air/water interface. The interstrand distance and the distance between two adjacent phenyl group was taken from ref. 42.

Molecular modelling of hairpin peptides assuming an ordered head-to-head β-hairpin packing

arrangement[42] suggests an occupied area of 259 Å^2 for one peptide (Scheme 2), which closely matches the limiting area of peptide **1b** on the water subphase. This result indicates that **1b** organizes in a head-to-head packing arrangement on the water surface as shown in Scheme 2, while a relatively smaller limiting area of **1a** suggests a less ordered peptide packing arrangement on the water surface.

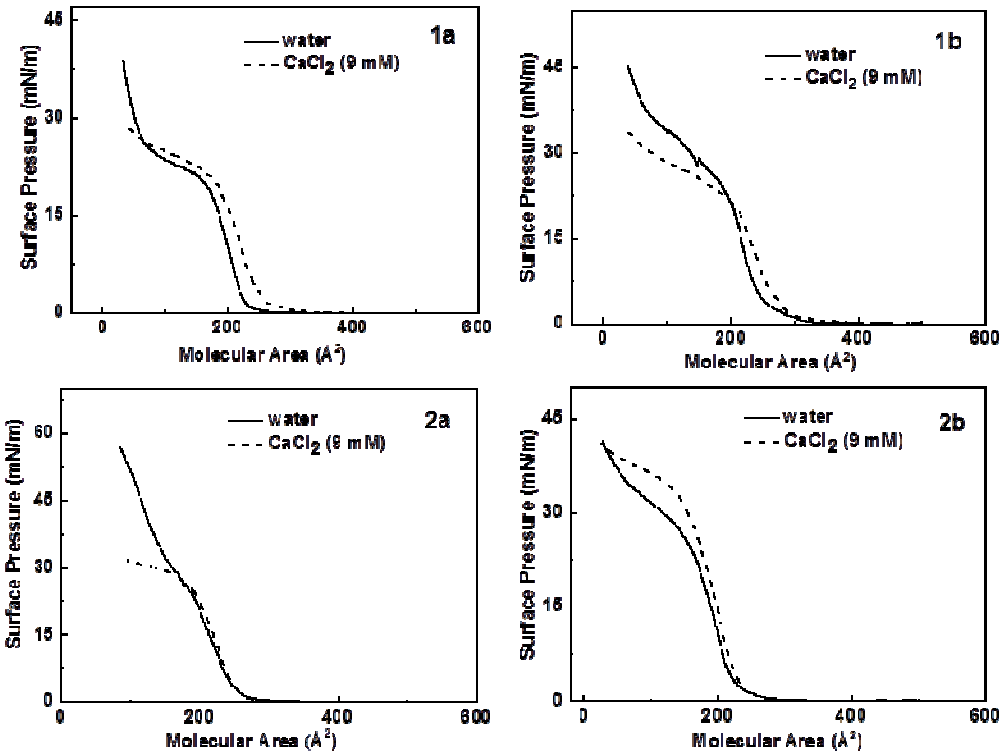


Figure 2. Surface pressure area isotherms of all hairpin peptides at 22°C. The solid lines indicate the isotherms on pure water surface. The dashed lines indicate the isotherms on aqueous $CaCl_2$ (9 mM) subphase.

The surface pressure increases after the phase transition of both peptides **1a** and **1b** on the water surface, and no obvious collapse pressure was observed within the range of the compression, demonstrating a strong flexibility of the peptide monolayers, i.e. the monolayers are highly compressible at very low surface area. In addition, the limiting area of the second liquid condensed phase is about half the area of the first liquid expanded limiting area for peptide **1a** and **1b** on the water surface, which indicates that the peptides overlap to form a double layer during the phase transition. For peptide **2a**, a significantly higher surface pressure was observed compared to **1a** or **1b** at the same molecular area, suggesting higher flexibility of peptide **2a** on the water surface. The isotherms of peptide **2b** have a shape similar to the isotherms of **1b** except for a smaller limiting area, which points at a tighter packing of **2b** on the water surface than **1b**.

On the subphase of $CaCl_2$, the isotherms of the peptides **1a** and **1b** both show significant shifts towards larger molecular area. The limiting areas both are about 30 Å^2 larger than the ones on

pure water surface, while the limiting areas for the peptides **2a** and **2b** change to a smaller value. The results indicate that the packing arrangement of the peptides on the Ca^{2+} ion subphase is closely related with the density of charged carboxyl groups.

The ordered head to head packing arrangement of peptide **1b** is considered to be a result of the subtle balance between hydrophobic interactions and carboxyl group interactions. However, on a subphase with calcium ions, this interaction balance is broken due to Ca^{2+} ion binding. Ion binding weakens intermolecular hydrogen bonding interactions and thus alters the inter- and intra-peptide distances. This explains the increase of the limiting areas of the peptides on Ca^{2+} ion subphase. Calculations that imply the modification of the distances between adjacent β-strands revealed that an interstrand distance change from 4.7 Å to 5.1 Å can explain the limiting area change of peptide **1b** on a $CaCl_2$ subphase.[43]

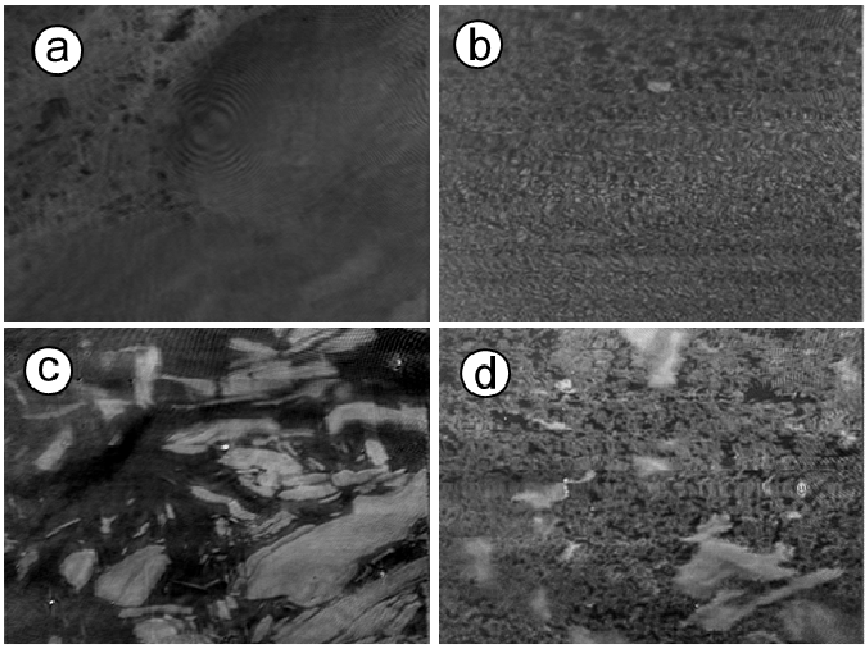


Figure 3. Brewster Angle Microscopy (BAM) images recorded for a monolayer of **1b** and **2b** spread on pure water and aqueous $CaCl_2$ (9 mM) subphase at the onset point of surface pressure. (a), **1b** on water subphase. (b), **1b** on $CaCl_2$ subphase. (c) **2b** on water subphase. (d), **2b** on $CaCl_2$ subphase. (The size of each BAM image corresponds to a monolayer area of 430 μm in width.)

BAM investigations of the peptides spread on water and on an aqueous $CaCl_2$ subphase were performed in order to elucidate monolayer properties in depth. Typical BAM images are shown for **1b** and **2b** in Figure 3, because the monolayer morphologies of **1a** and **1b** as well as the structures of **2a** and **2b** are similar both on water and $CaCl_2$ subphase. The BAM images of the peptides **1a** and **1b** on water show large domains with small holes inside (Figure 3a). Upon further compression, a uniform film was observed within the resolution of BAM (not shown). However, on the subphase of $CaCl_2$, the monolayer consists of many small domains at the onset pressure (Figure 3b). This monolayer morphology change suggests that the intermolecular interactions change in the monolayer

after binding with Ca^{2+} ions, which is consistent with the results of surface pressure-area isotherms.

Peptides **2a** and **2b** form larger domains on the water subphase at the onset of the surface pressure, while the film splits into many small domains on the subphase of $CaCl_2$. These domains merge and form a homogeneous monolayer at high pressure. The morphology difference of **2a** or **2b** on water and $CaCl_2$ subphase also suggests a strong binding interaction between Ca^{2+} ion and peptide monolayers.

Characterizations of peptide Langmuir-Blodgett films

The morphologies of all peptide Langmuir-Blodgett films (**1a**, **1b**, **2a** and **2b**) transferred from water to mica were imaged with AFM in tapping mode (Fig. 4 and supporting information). All four peptides form dense films on hydrophilic surfaces, each with a distinct morphology. Both peptides **1b** and **2b** organize in Langmuir films to give fibrous structures, while peptides **1a** and **2a** form connected clusters with random shapes.

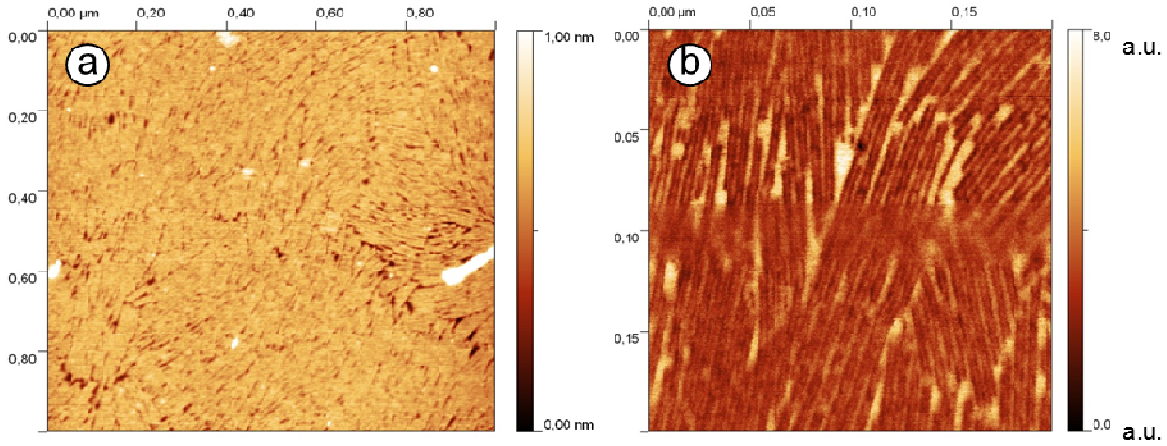


Figure 4. AFM images of peptide **1b** monolayer on mica transferred from a pure water subphase at π = 15 mN/m. (a) scanning area 1µm×1µm. (b) scanning area 0.2µm×0.2µm.

For the LB films with peptide **1b**, a distinctive arrangement was observed (Figure 4a). The molecules were seen to form highly regular stripe pattern with no preferential orientation, indicating lack of correlation between the peptide arrangement and the crystalline orientation of mica. Thus it can be excluded that the substrate itself has an impact on the structure of the peptide film. The measured film height is about 1 nm, confirming that only one monolayer has been transferred to the mica substrate. Phase images of the LB films (providing higher lateral resolution than the topographies) are shown in Figure 4b and reveal a double stripe pattern with a separation of 3.1nm±0.1nm. A few single strands are also observed. As the phase of an oscillating AFM tip depends on the specimen properties (such as elasticity and adhesion), the strong contrast in the AFM phase images may result from the tip interaction with hydrophobic and hydrophilic areas of the sample, or more specifically, with the peptide molecules (hydrophobic sidechains and hydrophilic

termini) and mica (in the film gaps) or partially adsorbed water molecules. The observed stripe pattern therefore confirms the assumption that the hairpins form β-sheets containing opposing rows of equally oriented molecules (each with a width of about 3 nm).

AFM scans were repeated on monolayers formed from different subphases from peptide **1b**. The LB films transferred from $CaCl_2$ (0.01 M) subphase exhibit shapes similar to the ones discussed above. The film height is 1.00 nm±0.03 nm, indicating that the subphase ions have not altered the dimensionality of the peptide film.

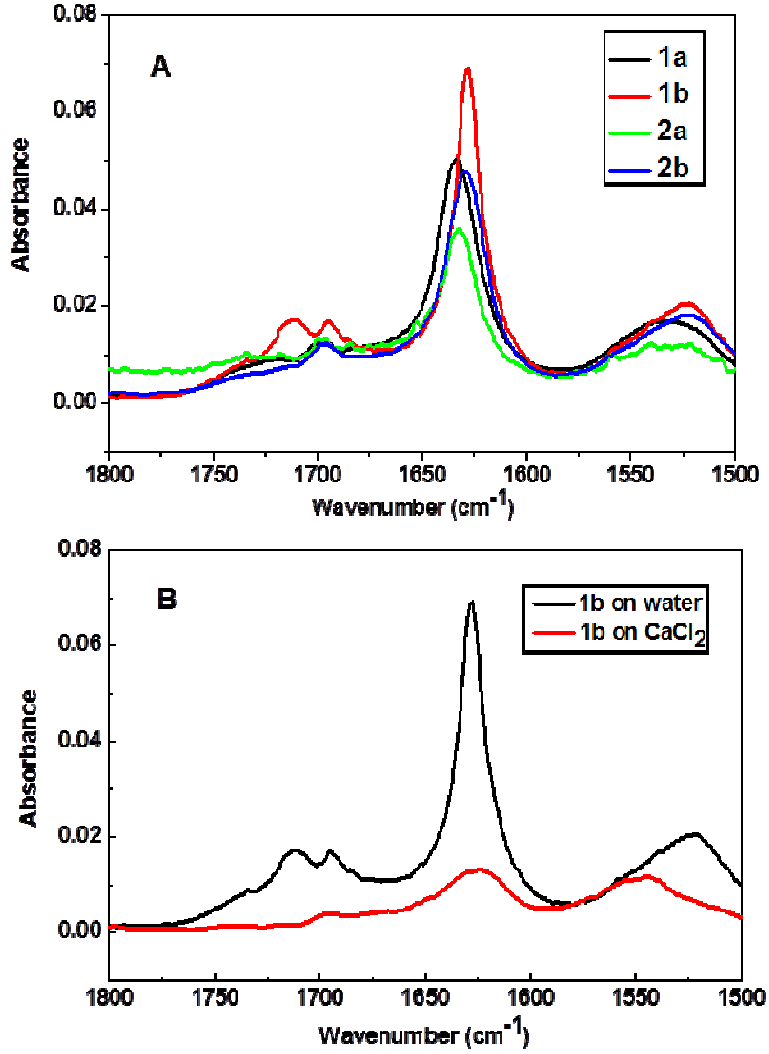


Figure 5. (A) FT-IR spectra of peptide LB multilayers transferred from water subphase at 15 mN/m. (B) FT-IR spectra of peptide **1b** multilayers transferred from water and $CaCl_2$ subphase (9 mM) at 15mN/m.

The secondary structures of the peptides were analyzed by FT-IR spectroscopy. Multilayers (i.e. 40 layers) of the peptides **1a** and **1b** were transferred onto CaF_2 substrates and analysed, as shown in Figure 5. Transfer ratios were recorded during multilayer transfer on CaF_2 substrates and they were found to be close to unity during all cycles and for all peptides being transferred. The IR absorption band between 1600-1700 cm^{-1} is assigned to the amide I stretching band, while the peak

position in this range is characteristic for the secondary structure of peptides. The amide I vibrational band around 1630 cm^{-1} and 1696 cm^{-1} existing in all peptide spectra indicate the presence of similar β-pleated sheet structure motifs of all peptides at the air/water interface.[44] FT-IR spectra were also recorded for peptide **1b** on a $CaCl_2$ subphase, as shown in Figure 5b. While the amide I band at 1630 cm^{-1} is still visible, the relative absorption intensity was significantly reduced due to the interactions with calcium ion stemming from the aqueous subphase.

Calcium carbonate crystallization beneath peptide monolayers

Figure 6 shows calcium carbonate crystals isolated from beneath a monolayer of peptide **1a** after 10 min to 4 hours growth time. In the first 10 minutes, olive-shaped particles with the size of 300-500 nm were observed by TEM (Figure 6a). They were very thin and could be easily decomposed by the high energy electron beam of the microscope. High Resolution TEM (HR-TEM) images showed that the particles are composed of amorphous calcium carbonate particles and contained single crystal domains in their interior (Figure 6c). After 1 hour of growth, the crystals dominating the surface were pancake-shaped (Figure 6b). Their Raman spectra showed a peak at 752 cm^{-1}, characteristic of vaterite polymorphs (see Figure S3, supporting information). After another 3 hours, two types of calcite crystals were isolated at the interface, as shown in Figure 6d-e. They were oriented with the {10.0} and {12.2} crystal faces pointing towards the monolayer. Remarkably, these orientations were only observed beneath **1a** peptide monolayers.

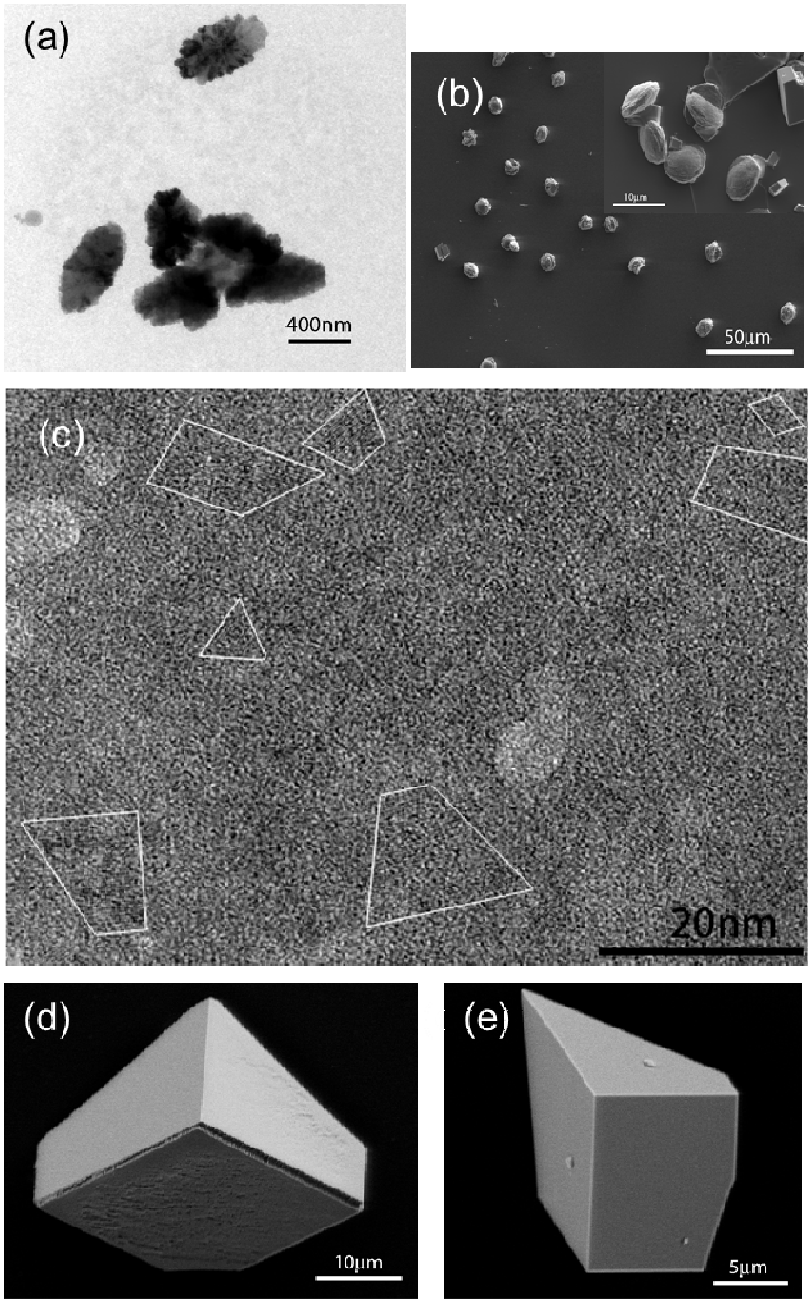


Figure 6. Crystallization of $CaCO_3$ beneath peptide **1a** monolayer at 0.1 mN/m. (a) TEM image of crystals separated after 10 min growth time. (b) SEM images of crystals separated after 1 hour. (c) High resolution TEM image of crystals separated after 10 min growth time. (d) and (e) SEM images of crystals separated after 4 hours.

Calcite crystals with a <01.2> orientation were dominating the surface for **1b** peptide monolayers after 4 hours of growth time (Figure 7d). The immature crystals isolated after 10 minutes growth time showed a triangular morphology (Figure 7a). Some particles had a rectangular morphology (Figure 7b) with inhomogeneous electron density distribution, indicating that these particles are composed of even smaller particles. HR-TEM showed that they are composed of amorphous calcium carbonate (ACC) with crystalline interior domains (Figure 7c). This indicates

that the crystals were formed via transformation from an amorphous phase into a crystalline one (Figure 7c).

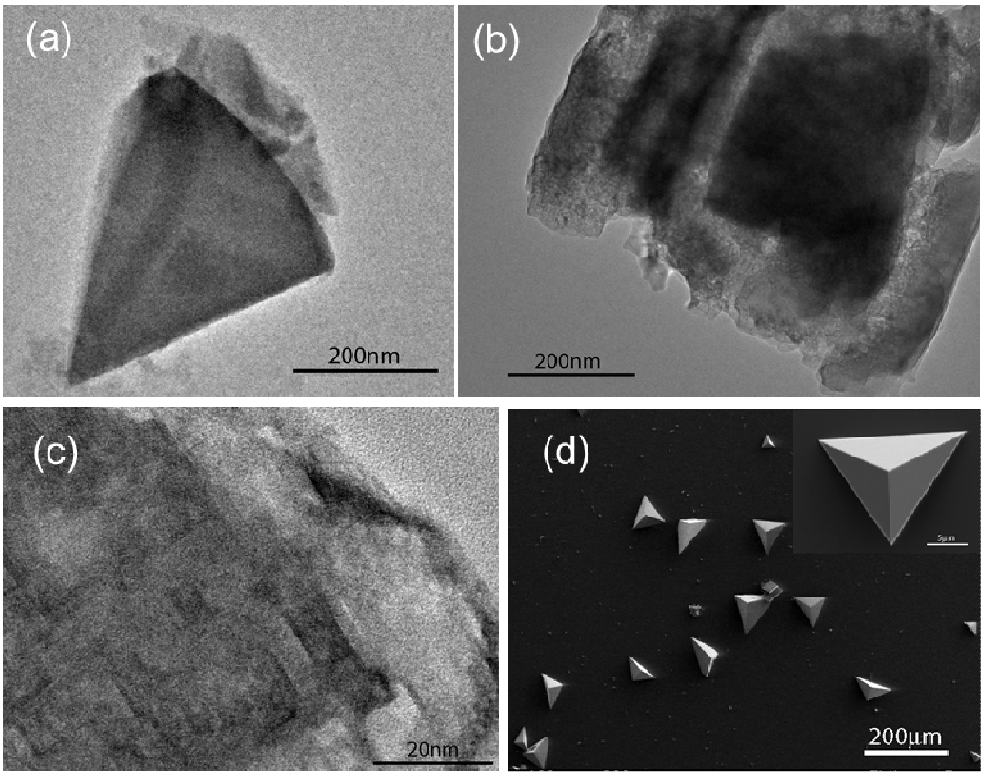


Figure 7. Crystallization of $CaCO_3$ beneath peptide **1a** monolayer at 0.1 mN/m. (a-c) TEM images of crystals separated after 10 min growth time. (d) SEM image of crystals separated after 4 hours.

Crystallization experiments beneath peptide **2a** and **2b** monolayers resulted in each case in {10.4} rhombohedral crystals with a concave pyramidal-shaped face. These crystals were predominant beneath the monolayer of peptide **2a**, but were less prominent underneath peptide **2b**. Here calcite crystals with <01.2> orientation could be additionally found (Figure 8).

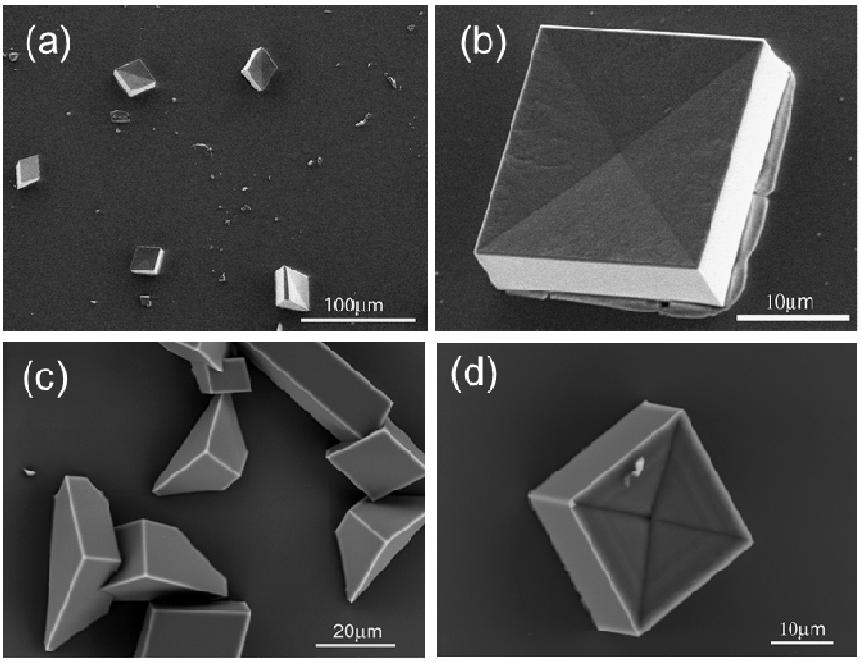


Figure 8. SEM images of $CaCO_3$ crystals grown beneath peptide **2a** (a,b) and **2b** (c, d) monolayer at 0.1 mN/m. Crystals were seperated after 4 hours (a,b) and 16 hours (c, d).

***In situ* observations of the crystallization process**

Observations of the crystallization were performed *in situ* employing optical and Brewster angle microscopy to further reveal the crystal growth process under peptide monolayers. These methods provide a direct and non-interfering way to observe crystal growth processes. Optical microscopy shows directly the crystal morphology and its distribution at the interface as well as the nucleation density, while BAM observations reveal monolayer domain changes during crystallization and $CaCO_3$ nucleation sites in the monolayer.

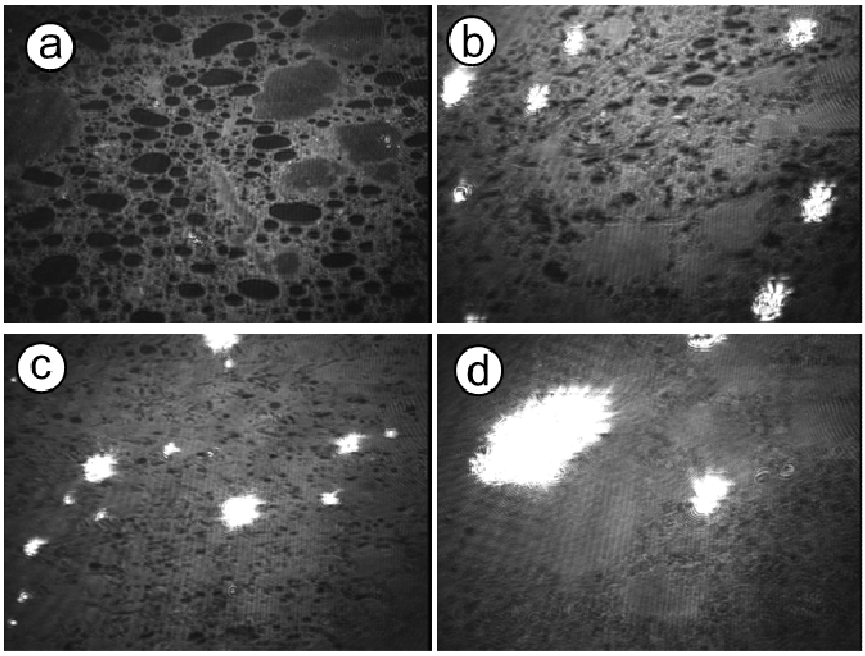


Figure 9. BAM images of $CaCO_3$ crystallization beneath peptide **1b** for 0, 1, 2, and 7 hours. The images were recorded at the onset point of surface pressure. (The size of each BAM image corresponds to a monolayer area of 430 μm in width.)

Immediately after spreading the **1b** peptide monolayer on $Ca(HCO_3)_2$ solution, domains with holes were observed with BAM (Figure 9). Bright spots appeared simultaneously, indicating calcium carbonate nucleation under these domains. The nucleation sites multiplied and grew with time, while the peptide domains became more and more tightly packed until they appeared completely homogenous (after approximately 7 h). Similar monolayer domain changes were also observed for other peptides during $CaCO_3$ crystallization (see Figs. S4 & S5 in the Supporting Information). As far as it is possible to judge from BAM images the growth of calcium carbonate seems to take place mainly at the edges of the peptide monolayer domains. Typical images of the crystallization process beneath **2b** monolayers are shown in Figure 10. Peptide **2b** forms irregular monolayer domains on an aqueous $Ca(HCO_3)_2$ solution right after spreading (Figure 10a). After a few minutes, some $CaCO_3$ particles could be seen at the edge of the domains (Figure 10b). Within 1 hour their number and size increased considerably. Simultaneously, the peptide domains were pushed together and formed a tightly packed monolayer within few hours of crystallization, as shown in Figure 10c and d.

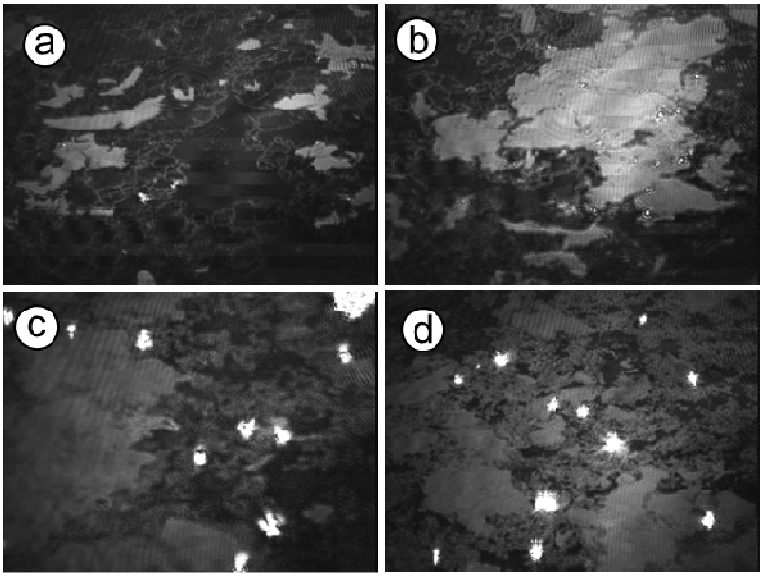
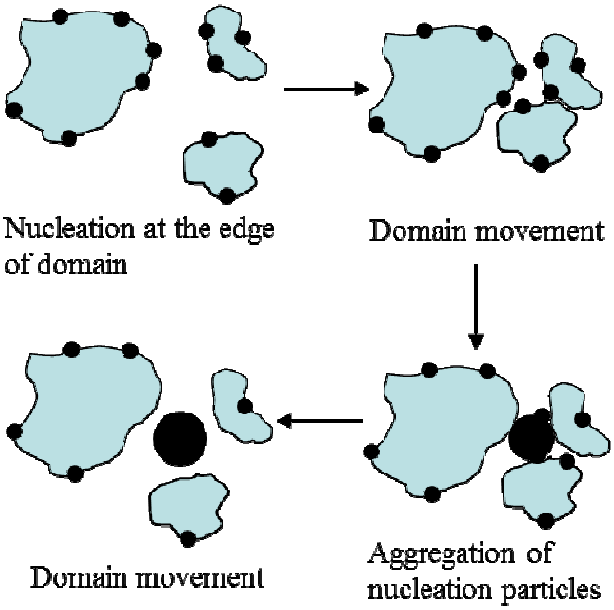


Figure 10. BAM images of $CaCO_3$ crystallization beneath peptide **2b** for 0, 1, 2, and 3 hours. The images was taken at the onset point of surface pressure. (The size of each BAM image corresponds to a monolayer area of 430 μm in width.)

Discussion.

Although Langmuir monolayers acting as $CaCO_3$ crystallization modifiers have been systematically investigated for over 20 years, many open questions still remain. For example, the hypothesis of a specific structure and/or geometric match between monolayers and certain crystal faces stands in contrast to the frequent occurrence of calcite crystals having their {01.2} crystal face oriented towards the monolayer, which is an experimental observation for a vast variety of structurally different monolayers.[45 49] The latter observation suggests that factors which are non-specific or characteristic of the monolayer structure govern $CaCO_3$ crystal growth and polymorph selection. In a few cases, the template model of selective monolayer-directed $CaCO_3$ crystallization seems to apply, which is the case for calcite crystals growing beneath amphiphilic monolayers of aliphatic sulfonates.[50] In this particular case the {00.1} crystal face of calcite seems to match the spatial arrangement of polar head groups in the monolayer.[51]

In the present study, different $CaCO_3$ crystal morphologies and crystal orientations were found to evolve underneath different hairpin peptide monolayers, suggesting that structural or geometrical factors of the monolayers might have some influence at different stages of crystal formation. As a feature common to all peptide monolayers we observed formation of an ACC precursor within the first minutes, thus pointing to a general crystallization pathway being operative under Langmuir monolayers at the early stage of crystal nucleation.



Scheme 3. Subsequent stages of calcium carbonate nucleation and aggregation, and changes of peptide monolayer domains during crystallization.

Scheme 3 highlights the likely succession of events accompanying the formation of $CaCO_3$ crystals beneath peptide monolayers. First, ACC primary particles (small black circles) form preferentially at the boundary of separate monolayer domains (blue patches), as indicated by time-resolved *in situ* BAM investigations. These nucleation sites might serve two different purposes: the edges of the monolayer are structurally flexible, being able to co-adapt to virtually any spatial arrangement of ions constituting the nascent ACC primary particles. Moreover, the non-covered water surface allows for efficient CO_2 escape from the aqueous subphase, leading to a local supersaturation of $CaCO_3$ at close proximity to the edges of the freely floating monolayer domains. The LC phase of the monolayer itself is characterized by a dense packing arrangement of amphiphilic molecules, and thus by a high surface charge density. The cumulative electrostatic effects should lead to preferential accumulation of Ca^{2+} and HCO_3^- ions underneath patches of the monolayer. However, the evaporation of CO_2 and thus the transformation of bicarbonate into carbonate should preferentially occur at the edges of monolayer domains, which might serve to rationalize the preferred formation of $CaCO_3$ particles at this location.

Where monolayer domains approach each other, larger aggregates of ACC particles can thus form via fusion of smaller particles being attached to edges of the monolayer domains (Scheme 3). According to results from TEM investigations $CaCO_3$ crystals start to grow from the inside of larger ACC particles, giving rise to $CaCO_3$ single crystals. This statement gleans (indirect) support from the fact that orientations, sizes and shapes of $CaCO_3$ crystals observed in the present and similar model systems often show broad distributions, which lack a specific and direct correlation with regard to the structure of the templating monolayer.

ACC phases have been observed in natural nacre[52] and artificial model systems.[5359] It is reported that these ACC phases are mostly stabilized by polyanions such as poly(aspartic acid)[60] or poly(acrylic acid)[61] in solution. It is thus reasonable to assume that a thin layer of ACC can also form beneath monolayers of acidic peptides owing to their polyanionic character. Hence, $CaCO_3$ crystallization starts subsequent to the accumulation of amorphous particles, and there is no experimental evidence for a direct epitaxy of $CaCO_3$ crystals emanating from a two-dimensional patch of the peptide monolayer. Our findings seem to support recent studies of multistep crystallization which in a colloidal model system demonstrated the formation of crystal nuclei of subcritical size inside an amorphous precursor phase and their subsequent fusion and transformation into single crystals.[62]

Subsequent to the formation of ACC particles, vaterite crystals have been observed in the early stage of all crystallization experiments performed with peptide monolayers. After 4 hours the proportion of vaterite crystals decreased significantly for peptide **1a** monolayers, and the number of calcite crystals increased at the same time. We conclude that vaterite is a transient polymorph which transforms into calcite during the later stage of the crystallization process. The relatively slow transformation of vaterite into calcite suggests, however, that this process might be influenced by the monolayer. However, for peptide **1b** monolayers, vaterite and calcite crystals occurred simultaneously and their relative ratio did not change significantly within the experimental period. Since a less ordered β-hairpin packing arrangement for peptide **1a** is indicated by surface pressure-area isotherms, FTIR and AFM investigations, we might tentatively suggest that the more flexible nature of peptide **1a** compared with peptide **1b**, leads to rapid formation of vaterite nuclei under **1a** monolayers and their subsequent transformation into <10.0> oriented calcite crystals. Similar results have been suggested by Kuther et al.[63] for a self-assembled monolayer showing a low degree of structural organisation, which promoted the formation of aragonite over calcite under experimental conditions different to ours.

Various orientations and morphologies of calcite crystals were observed for the different peptide monolayers at the late stages of crystallization, i.e. after a few hours. For peptides **1a** and **1b**, XPS results show that individual peptide molecules on average bind about 4 calcium ions if being spread on an aqueous subphase containing calcium acetate (supporting information S6). Therefore, the different orientations of calcite crystals grown under **1a** and **1b** peptide monolayers cannot be attributed to different monolayer charge densities. AFM investigations indicate that **1b** assumes a highly ordered head-to-head packing arrangement (cf. Scheme 2), while **1a** self-assembles into a less well-organized packing arrangement. Based on these results, we conclude that monolayer structure flexibility, i.e. the ability of the monolayer to co-adapt to structural changes imposed by the aggregating and transforming $CaCO_3$ particles is a crucial factor in controlling the orientations and shapes of the mature calcite single crystals.

Conclusions

The monolayers of four acidic model hairpin peptides were investigated for templating effects with respect to $CaCO_3$ crystallization. Unlike in solution, all peptides adopt a β-hairpin conformation at the air/water interface, which results from the sequence of alternating hydrophilic and hydrophobic amino acid residues. Peptide **1b** forms a highly ordered head-to-head packing arrangement in the monolayer, whereas peptide **1a** is structurally more flexible. These slight structural differences cause a significant variation of $CaCO_3$ crystals that grow underneath the monolayers. The crystals were characterized by SEM, TEM, SAED, and *in situ* optical and BAM. An amorphous calcium carbonate precursor was found during the initial stage of crystallization. TEM images indicate that this amorphous precursor contains nanosized crystals which form preferentially inside the amorphous phase, indicating that the crystallization process starts subsequent to aggregation and merging of amorphous precursor particles. The formation of amorphous calcium carbonate particles mainly occurs at the boundary of LC domains or in the LE phase of the peptide monolayers, the movement of monolayer domains being crucial for the lateral diffusion or aggregation of ACC particles. $CaCO_3$ crystals preferentially form within fused patches of ACC particles. It should be noted that in our hands all attempts have failed to provide experimental evidence for a direct and two-dimensional contact between the $CaCO_3$ crystals and the peptide monolayer domains. This statement stands in sharp contrast to the original notion of a templating monolayer and the mechanistic concept of heteroepitaxial crystal growth in such model systems. The significance of ACC precursor phases yet remains an open question. It seems evident that the ACC phase serves as nutrient solution for the maturing $CaCO_3$ crystals. If and how the ACC phase influences the growth characteristics of single crystals emanating from its surrounding precursor currently remains an open question, as does the functional role of the monolayer during this stage of crystal growth. Addressing such fundamental aspects of biologically inspired crystallisation will require experimental set-ups different to those employed in the present investigations.

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Supporting Information Available

Additional AFM topography images of Langmuir-Blodgett films of peptide monolayer transferred from water subphase, BAM images of $CaCO_3$ crystallization under monolayers of peptide **1a** and **2a**. Raman spectrum of crystals grown under the monolayer of peptide **1a**. This material is available free of charge via the Internet at www.pubs.acs.org.

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