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Bronze, silver and gold: functionalized group 11 organotin sulfide clusters†

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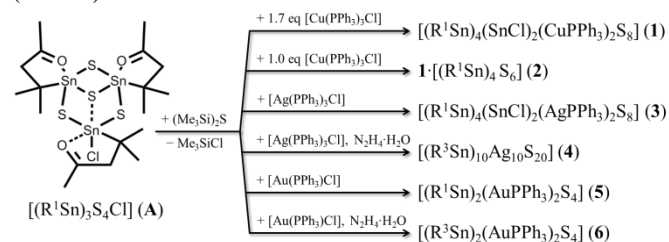
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Synthesis, properties and reactivity of group 11 organotin sulfide clusters $[(R^1Sn)_4(SnCl)_2(MPPh_3)_2S_8]$ ($M = Cu, Ag$), $[(R^3Sn)_{10}Ag_{10}S_{20}]$, and $[(R^{1,3}Sn)_2(AuPPh_3)_2S_4]$ with covalently bound, carbonyl or hydrazine-terminated ligands $R^1 = CMe_2CH_2C(Me)O$ or $R^3 = CMe_2CH_2C(Me)NNH_2$ are reported.

The last decades afforded a large variety of ligand-protected chalcogenide clusters of group 11 metals.¹ Also ternary complexes and clusters were reported,² which often combined structural and physical properties of the binary components.³ Additionally, with regard to further reactivity, the introduction of functionalized phosphines was achieved.⁴ Consequently, the combination of functional organic ligand shells with finely tunable (ternary or multinary) inorganic cores, providing even more diversity, is thus of interest for diverse purposes, such as for optoelectronics or solar cell development.⁵ To address according functional materials, the synthesis and directed derivatization of clusters with covalently attached organic groups is currently investigated in detail for polyoxometalates⁶ as well as for group 14/16 compounds,⁷ for instance.

The organic groups on the periphery of organogermanium and organotin sulfide complexes of the general type $[(R^T)_xS_y]$ ($R^T =$ functionalized organic ligand: $R^1 = CMe_2CH_2C(Me)O$, $HR^2 = C_2H_4COOH$; $T = Ge, Sn$; $x/y = 4/6, 6/10$) are reactive towards hydrazine derivatives, which allowed further derivatizations of the organic ligand shell.⁸ Hereby, an extension of the ligand shell towards ligands such as $R^3 = CMe_2CH_2C(Me)NNH_2$, $R^4 = \{CMe_2CH_2C(Me)NNH\}_2CO$, or $R^5 = CMe_2CH_2C(Me)NNC(2-py)_2$ was achieved. Reactions with transition metal compounds additionally enabled the extension of the inorganic core to form the organofunctionalized ternary clusters $[(R^3T)_2(CuPR_3)_6S_6]$,⁸ $[(R^2Sn)_6(OMe)_6Cu_2S_6]^{4-}$,⁸ $[(R^3Ge)_4Pd_6S_{12}]$ ⁹ and $[(R^4)_2(Sn(\mu-S)_2Cu-PPH_2Me)_4]$.¹⁰ Bispyridine-decorated clusters $[(R^5Sn)_4(ZnX)_4S_8]$ ($X = Cl, Br, I$) were obtained by slow diffusion of zinc halide solutions into solutions of the organotin sulfide cluster $[(R^5T)_6S_{10}]$.¹¹ Further structurally characterized ternary group 11 metal/organotin sulfide clusters have so far been surrounded by unreactive organic groups, as in $[(PhSn)_2(CuPPh_2Me)_6S_6]$,¹² $[(PhSn)_{12}Cu_{19}(PEt_2Ph)_3S_{28}]^+$,¹³ and the mixed-valence compound $[(CH_2)_4Sn^{IV}]_6(CuPPh_3)_6Sn^{II}Cu_4S_{12}]$.¹⁴ Ternary complexes containing Ag/Sn/S or Au/Sn/S cores have not been characterized yet by X-ray diffraction.

Recently we reported new functional organotin sulfide complexes of the type $[(R^{1,3}Sn)_3S_4Cl]$ that exhibit defect heterocubane scaffolds with carbonyl and hydrazone groups.¹⁵ Herein, we show that these turned out to be suitable precursor complexes, as well. $[(R^1Sn)_3S_4Cl]$ (**A**) reacts with the group 11 metal complexes $[Cu(PPh_3)_3Cl]$, $[Ag(PPh_3)_3Cl]$ and $[Au(PPh_3)Cl]$ in CH_2Cl_2 with $(Me_3Si)_2S$, in the cases of Ag and Au complexes with subsequent *in situ* derivatization by hydrazine hydrate (Scheme 1), to yield single crystals of the following compounds comprising ternary clusters $[(R^1Sn)_4(SnCl)_2(CuPPh_3)_2S_8] \cdot 4 CH_2Cl_2$ (**1**·4 CH_2Cl_2), **1**· $[(R^1Sn)_4S_6]$ (hereafter denoted as **2**), $[(R^1Sn)_4(SnCl)_2(CuPPh_3)_2S_8] \cdot [(R^1Sn)_4S_6]$ (**3**· $[(R^1Sn)_4S_6]$), $[(R^3Sn)_{10}Ag_{10}S_{20}] \cdot 3.5CH_2Cl_2$ (**4**·3.5 CH_2Cl_2), $[(R^1Sn(\mu-S))_2\{AuPPh_3(\mu-S)\}_2]$ (**5**), and $[(R^3Sn(\mu-S))_2\{AuPPh_3(\mu-S)\}_2] \cdot 5CH_2Cl_2$ (**6**·5 CH_2Cl_2). The clusters exhibit different topologies and compositions and comprise reactive substituents with either carbonyl or hydrazone groups. In addition we report the single crystal X-ray diffraction analysis of $(Me_3Si)_2S$ (see ESI).



Scheme 1. Synthesis of compounds **1–6** ($R^1 = CMe_2CH_2C(Me)O$, $R^3 = CMe_2CH_2C(Me)NNH_2$).

The reaction of **A** with 1.7 eq $[Cu(PPh_3)_3Cl]$ and $(Me_3Si)_2S$ in CH_2Cl_2 , and subsequent layering with *n*-hexane yielded $[(R^1Sn)_4(SnCl)_2(CuPPh_3)_2S_8] \cdot 4 CH_2Cl_2$ (**1**·4 CH_2Cl_2) as light yellow crystals. In the molecular structure of the cluster, two $\{(R^1Sn(\mu-S)(\mu_3-S))_2\}$ moieties are linked by two $\{(CuPPh_3)(SnCl)\}$ units (see Fig. 1). The R^1Sn atoms Sn1 and Sn2 show a trigonal bipyramidal coordination environment, with the axial positions being occupied by O1, O2 and S1, whereas S2, S3 and C1 (Sn1) or S2, S4^a and C7 (Sn2) are situated in the equatorial positions. The μ_3 -S-bridging atoms S1 and S4, and the μ -bridging atom S3 connect the $\{(R^1Sn(\mu-S)(\mu_3-S))_2\}$ -moieties to the central part of the cluster. Here, Cu1 is surrounded in a tetrahedral fashion by a PPh_3 ligand,

S1, S4 and Sn3. The latter has a formal oxidation state of +II and exhibits a distorted *pseudo*-trigonal bipyramidal environment, with Cu1 and S3^a in the “axial” positions (Cu1–Sn3–S3^a 149.75(2)°), and Cl1, S4^a and the sterically active lone pair in the “equatorial” positions; further *pseudo*-trigonal bipyramidal coordination geometry of Sn^{II} atoms have been described by *Jurkschat* and co-workers.¹⁶ The mixed-valence situation in **1** was confirmed by DFT calculations (see below). As another peculiarity, **1** comprises a rare Cu–Sn bond (Cu1–Sn3 2.6054(4) Å) which was formed *in situ*.¹⁷ It is not possible to monitor the formation process in detail due to the fairly complex reaction mixture and due to only poor solubility of the product. However, we assume that the reduction of RSn^{IV} to Sn^{II} comes along with the oxidation of the released organic ligand as discussed for the mixed valence compounds [(R^{1,3}Sn)₃Se₄][SnCl₃] under similar reaction conditions;¹⁵ the oxidation of PPh₃ or Cu^I, however, was not observed under the reaction conditions, which was clarified by NMR studies and test for Cu²⁺ with NH₃.

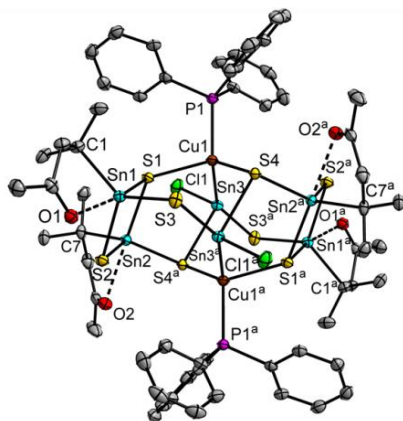


Figure 1. Molecular structure of **1** without H atoms (ellipsoids drawn at 50% probability); a = 1–x, –y, –z. Selected structural parameter [Å, °]: Sn–S 2.3654(8)–2.5058(8), Sn(1,2)–C 2.183(3)–2.192(3), Sn(1,2)–O(1,2) 2.516(2)–2.537(2), Sn3–Cl1 2.4713(8), Sn3–Cu1 2.6054(4), Cu1–S(1,4) 2.3626(9)–2.3166(8), Cu1–P1 2.2590(8); Sn–S–Sn 86.57(2)–89.70(3), S–Sn–S 91.24(3)–121.81(3), Cu1–Sn3–S3^a 149.75(2), Cu1–Sn3–S4^a 107.60(2), Cl1–Sn3–Cu1 111.23(2), Cl1–Sn3–S3^a 90.20(3), Cl1–Sn3–S4^a 95.70(3), C–Sn–O 72.10(10)–73.71(10).

The observation of formal Sn^{II} and Sn^{IV} atoms is in agreement with natural charges, obtained by natural population analyses (NPA)¹⁸ of the DFT wave function, which was calculated by simultaneous optimization of the geometric and electronic structure using the program system TURBOMOLE.¹⁹ The formal Sn^{IV} atoms exhibit charges of +1.38 and +1.47, which are by a factor of 1.6 to 1.7 larger than the +0.88 charge calculated for the formal Sn^{II} atoms. Details of the DFT calculations are provided in the ESI.

Neither variation of the stoichiometry of the starting materials nor attempts to replace the Cl[–] ligand have led to the isolation of further clusters until now. An analogous reaction with 1.0 eq [Cu(PPh₃)₃Cl], however, yielded **1**:[(R¹Sn)₄S₆] (**2**). The structural parameters of the two large co-crystallizing clusters in **2** are similar to those observed in **1**:4 CH₂Cl₂ (see ESI) and [(R¹Sn)₄S₆].²⁰ Upon according reaction with [Ag(PPh₃)₃Cl] and (Me₃Si)₂S at –78 °C in CH₂Cl₂, and subsequent, careful warming-up to room temperature, a colorless precipitate was obtained, which was re-dissolved by addition of excess CH₂Cl₂. Layering with *n*-hexane afforded [(R¹Sn)₄(SnCl)₂(AgPPh₃)₂S₈][[(R¹Sn)₄S₆] (3:[(R¹Sn)₄S₆]) as light yellow crystals. The ternary cluster is isostructural to that in **2**, hence the structural parameters are similar to those observed in **1** or **2**, except the expected elongation of M–S, M–P and M–Sn bonds (Ag1–Sn3 2.6803(3) Å) for M = Ag in comparison with M = Cu, and

a larger Sn–M–P angle (**1**: 115.40(3) for M = Cu, **3**: 131.23(2) for M = Ag). For selected structural parameters and NPA results see ESI.

Addition of hydrazine hydrate to the re-dissolved solution, and subsequent layering with *n*-hexane produced orange crystals of the hydrazone functionalized cluster [(R³Sn)₁₀Ag₁₀S₂₀]:3.5 CH₂Cl₂ (4:3.5 CH₂Cl₂). **4** consists of three {R³Sn(μ-S)(μ₃-S)}₂ units and two {[(R³Sn(μ-S))₂(μ₃-S)(μ₄-S)]} moieties that are linked by ten Ag atoms (see Fig. 2). All Sn atoms exhibit trigonal bipyramidal coordination, with N and S atoms in the axial positions. Four Ag atoms (Ag1, Ag5, Ag7, Ag10) are coordinated in a linear manner, while six other Ag atoms exhibit trigonal planar coordination. Closest Ag⁺–Ag distances are in the range of 2.9353(7) Å to 3.3260(7) Å. Eight of the Ag atoms are connected this way, with a central butterfly-like arrangement and to adjacent Ag⁺–Ag units, while one Ag⁺–Ag unit is further apart. The cluster possesses *pseudo*-C₂-symmetry. The *pseudo*-C₂-axis runs through the centers of the Ag⁺–Ag four-ring and the separate Ag⁺–Ag dimer. In contrast to known silver sulfide clusters,^{1a} the cluster does not additionally contain stabilizing phosphine groups attached to the silver atoms. N→Sn coordination of the bidentate organic ligands seems to provide sufficient kinetic stabilization.

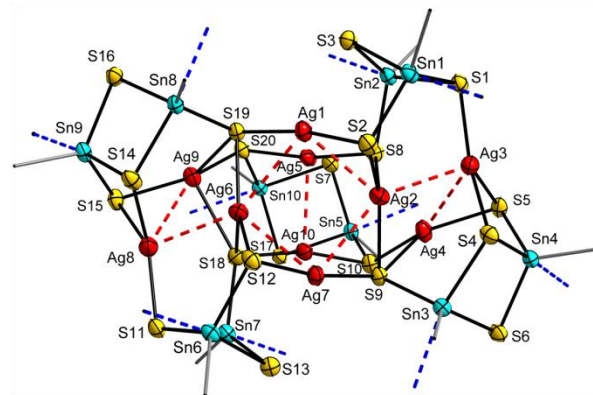


Figure 2. Molecular structure of the inorganic core of **4** along the *pseudo*-C₂-axis (ellipsoids drawn at 50% probability, bonds to organic substituents are indicated by dashed lines). Ag⁺–Ag contacts below 3.4 Å are drawn as dashed red lines. Selected structural parameters [Å, °]: Sn–S 2.3865(16)–2.5987(17), Sn–C 2.170(6)–2.187(7), Sn[–]N 2.362(5)–2.456(5), Ag–S 2.3666(16)–2.9040(16), Ag⁺–Ag 2.9353(7)–3.3260(7); S–Sn–S 89.57(5)–122.29(6), S–Ag–S 83.47(5)–169.97(6), Sn–S–Sn 85.13(5)–89.69(5), Ag–S–Ag, 72.65(5)–146.30(6), Ag–S–Sn 80.15(5)–144.91(7), C–Sn–N 71.6(2)–75.8(2).

The employment of a related gold complex, [Au(PPh₃)Cl], along with (Me₃Si)₂S in the reaction with **A** afforded single-crystals of the complex [{(R¹Sn(μ-S))₂}{Au(PPh₃(μ-S))₂}] (**5**). Moreover, addition of hydrazine hydrate to the reaction mixture yielded the hydrazone functionalized derivative [{(R³Sn(μ-S))₂}{Au(PPh₃(μ-S))₂}]:5 CH₂Cl₂ (**6**:5 CH₂Cl₂) (see Figure 3). In the present case, the derivatization of the organic ligand did not affect the molecular structure of the complex. In **5** and **6**, a central {R^{1,3}Sn(μ-S)}₂ unit (*c.f.* ref. [21]) is terminated by two {Au(PPh₃(μ-S))} groups. As in **1–4**, organo-decorated Sn atoms exhibit a trigonal bipyramidal coordination environment, while all Au atoms show linear coordination. The structural parameters of the central part is similar in both complexes, but the orientation of the attached {Au(PPh₃(μ-S))} units is different. In **5**, they point away from the central Sn₂S₂ ring, with a C1–Sn1–S2–Au1 *trans* arrangement, whereas in **6**, the according atoms show a *cis* arrangement. While interatomic distances are similar in both complexes, the Sn1–S2–Au1 and S2–Au1–P1 angles show notable differences in both complexes, indicating the influence of the functional group included in the organic ligand.

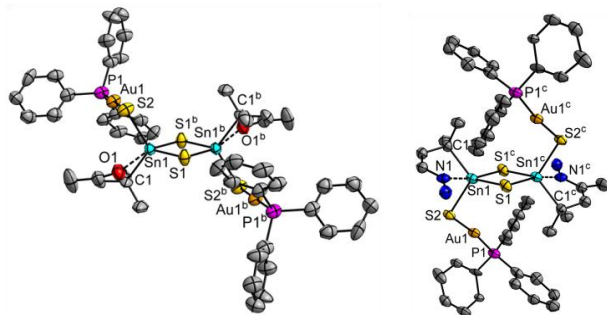


Figure 3. Molecular structures of **5** (left) and **6** (right) without H atoms (ellipsoids drawn at 50% probability). $b = -x, 2-y, 1-z$; $c = 2-x, 1-y, -z$. Selected structural parameters [Å, °]: **5**: Sn–S 2.380(2)–2.463(2), Sn1–C1 2.189(8), Sn1–O1 2.714(6), Au1–S2 2.291(2), Au1–P1 2.260(2); Sn1–S1–Sn1^b 86.30(7), S1–Sn1–S1^b 93.70(7), Sn1–S2–Au1 96.63(9), S2–Au1–P1 178.29(9); **6**: Sn–S 2.389(3)–2.525(3), Sn1–C1 2.186(14), Sn1–N1 2.479(11), Au1–S2 2.298(3), Au1–P1 2.250(3); Sn1–S1–Sn1^c 89.06(11), S1–Sn1–S1^c 90.94(11), Sn1–S2–Au1 103.86(13), S2–Au1–P1 172.84(12).

In summary, a series of ternary group 11 organotin sulfide clusters have been synthesized and structurally characterized. The choice of the very group 11 metal M – although added as similar complexes [M(PPh₃)_xCl] (M/x = Cu/3, Ag/3, Au/1) – causes specific coordination environments that affect the molecular structures. Accordingly, the products exhibit different compositions and topologies. Unexpectedly, the reaction with the Cu and Ag precursor induced *in situ* reduction of one third of the Sn atoms under Sn–Cu and Sn–Ag bond formation. It was shown for the first time that functionalized ternary M/Sn/S complex can undergo further derivatization. This work was supported by the Deutsche Forschungsgemeinschaft within the framework of GRK1782.

Notes and references

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† X-ray crystallographic data: Data collection on a STOE IPDS2 diffractometer using graphite-monochromatized Mo K_α radiation ($\lambda = 0.71073$ Å) at 100 K. Structure solution and refinement by direct methods and fullmatrix least-squares on F², respectively; SHELXTL software.²²

Electronic Supplementary Information (ESI) available: Further details of syntheses, analyses, and crystallographic data/further figures for **1–6** and (Me₃Si)₂S (CCDC 1011096–1011102). See DOI: 10.1039/c000000x/

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