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COMMUNICATION

High-Throughput Colorimetric Assays for Mercury (II) in Blood and Wastewater Based on Mercury-Stimulated Catalysis Activities of Small Silver Nanoparticles in Temperature-Switchable Gelatin Matrix

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A catalysis-based, label-free, and high-throughput colorimetric protocol has been initially proposed for detecting mercury (II) in blood and wastewater with 96-cell plates, based on mercury-enhanced catalysis activity of small silver nanoparticles synthesized in gelatin matrix with unique temperature switchable sol-gel transition.

Recent years have witnessed the rapid development of nano-scaled materials of noble metals (i.e., Au, Ag, Pt, and Pd) with many fantastic physicochemical properties¹. Especially, their unique enzyme-like catalysis has attracted considerable interests in catalysis and biochemical analysis applications. For example, gold nanoparticles (NPs) could present peroxidase-like catalysis activity for sensing H₂O₂ and glucose². Catalytic Pt NPs have also been utilized for the detections of H₂O₂ and scavenging superoxide free radicals². Moreover, the catalytic activities of noble metal nanomaterials can substantially depend on their sizes, known as the “size effect”, for example, high catalytic activity can occur for gold NPs with sizes smaller than 5.0 nm³. As a result, many efforts have been devoted to the synthesis of size-reducing gold NPs or nanoclusters⁴. The preparation of cheap and catalytic small silver NPs (AgNPs), however, still remains a challenging but abstractive issue, due to the difficult synthesis, low stability, and poor catalysis performances of small AgNPs.

As the most hazardous heavy metal ions, mercury ions were targeted by many modern detection methods typically as the DNA folding-based electrochemical detection, the fluorescence quenching-based analysis, and gold or silver NPs-based colorimetric assay⁵. However, these methods might suffer from either of low detection sensitivity and throughput, or poor analysis stability and abilities against background interferences. Recently, some catalytic noble metal nanomaterials have been utilized for probing heavy metal ions like Hg²⁺ ions that could inhibit and stimulate their catalysis activities through the specific Au⁺-Hg²⁺ interaction and the change of the surface properties of AuNPs or AgNPs by forming Ag-Hg alloys, respectively⁶. Such a catalysis-based analysis methodology may possess some advantages over the traditional methods like fluorescence-based ones in terms of the detection stability and targeting abilities against interferences from the complicated media (i.e., blood).

Gelatin, as a natural biopolymer derived from collagen widely used in food products and medicines⁷, is nontoxic, non-

immunogenic, biodegradable, and especially with unique sol-gel transition (critical transition at 35 °C). Such a biomaterial was introduced for the synthesis of AgNPs but entailing additional reductants like maltose and citrate⁸. Also, rare success has been reported for small AgNPs with Hg²⁺-enhanced catalysis activity. In this work, gelatin matrix was employed acting as the reducing agent and stabilizing matrix for the one-pot synthesis of small AgNPs under physiological temperature (37 °C) without adding any reducing agents. Unexpectedly, the prepared Gel-AgNPs could display Hg²⁺-stimulated powerful catalysis activities in catalyzing 3,3',5,5'-tetramethylbenzidine (TMB)-H₂O₂ reactions.

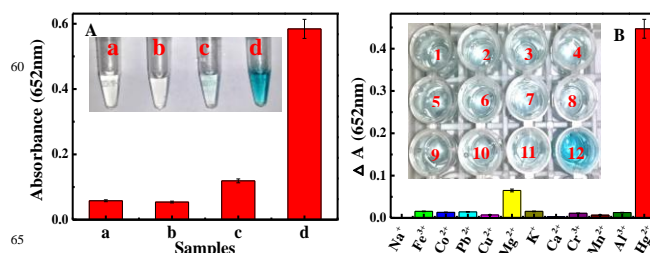
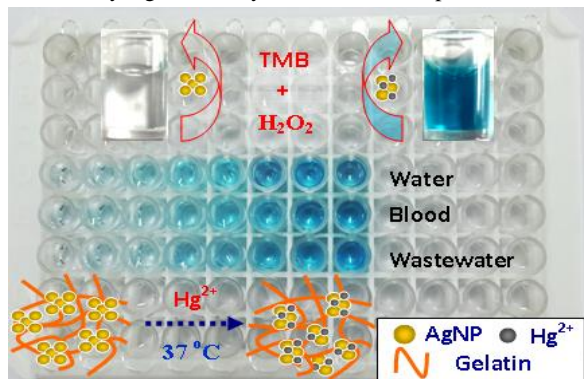


Fig. 1 (A) Comparison of catalysis activities in TMB-H₂O₂ reactions among (a) water (control), (b) 8.0 × 10⁻⁷ M Hg²⁺, (c) 7.2 × 10⁻⁶ M Gel-AgNPs, and (d) 8.0 × 10⁻⁷ M Hg²⁺ with 7.2 × 10⁻⁶ M Gel-AgNPs, corresponding to colorimetric photographs (Insert); (B) Investigation of catalytic activities of Gel-AgNPs in TMB-H₂O₂ reactions stimulated by various metal ions of Na⁺, Fe³⁺, Co²⁺, Pb²⁺, Cu²⁺, Mg²⁺, K⁺, Ca²⁺, Cr³⁺, Mn²⁺, Al³⁺ (each 5.0 × 10⁻⁵ M), and Hg²⁺ (4.5 × 10⁻⁷ M), with corresponding colorimetric photographs (Insert, sample 1–12).

The special phenomenon of mercury (II)-enhanced catalysis of small Gel-AgNPs was explored by comparable tests in the absence and presence of Hg²⁺ ions (**Fig. 1A**). It was found that Gel-AgNPs showed a little of catalysis in the TMB-H₂O₂ reaction at the meaningfully low concentration (**Fig. 1A** (c)). However, their catalysis activities could be significantly enhanced by Hg²⁺ ions, showing deep blue reaction product in the photograph (**Fig. 1A** (d)). Furthermore, effects of some other ions on stimulating the catalytic activities of Gel-AgNPs were investigated (**Fig. 1B**, and **Fig. S1**, ESI). One can observe that common inorganic ions presented no capabilities of stimulating the catalytic activities of Gel-AgNPs, even though at concentrations about 100-fold higher than that of Hg²⁺ ions. Based on the specific mercury-stimulated catalysis activity of Gel-AgNPs, a label-free colorimetric protocol has been thus proposed for the detection of Hg²⁺ ions. **Scheme 1**

illustrates the catalysis-enhanced analysis protocol using 96-well plates towards high-throughput colorimetric assays for Hg^{2+} ions in water, blood, and wastewater. As described in **Scheme 1**, Gel-AgNPs (thawed at $37\text{ }^\circ\text{C}$) alone (left top) exhibited no significant catalysis in TMB- H_2O_2 reactions at the low concentration. In contrast, the catalysis activity of Gel-AgNPs could be greatly enhanced by Hg^{2+} ions to yield blue reaction products.



Scheme 1 Colorimetric assays of Hg^{2+} in different samples by the TMB- H_2O_2 reactions catalyzed by Gel-AgNPs with Hg^{2+} -enhanced catalysis activities, where different concentrations of Hg^{2+} ions in water, blood, and wastewater were measured using 96-well plates with Gel-AgNPs.

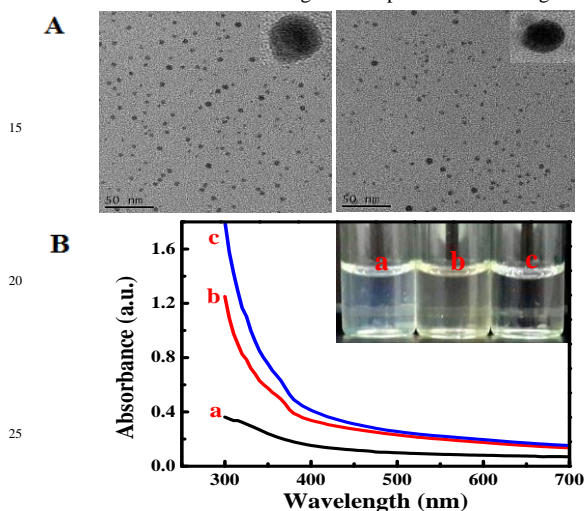


Fig. 2 Characterization of (A) TEM images for Gel-AgNPs alone (left) and Gel-AgNPs with Hg^{2+} (right), each with amplified view of one particle (Insert), and (B) UV-vis spectra of (a) Gel, (b) Gel-AgNPs alone, and (c) Gel-AgNPs with Hg^{2+} , corresponding to the photographs (Insert).

Transmission electron microscopy (TEM) and UV-vis spectrophotometer were separately utilized to characterize Gel-AgNPs in the absence and presence of mercury (II) (**Fig. 2**). As shown in **Fig. 2A**, original Gel-AgNPs were dense and well dispersed with the average particle diameter of about 5.0 nm (**Fig. 2A**, left). When Hg^{2+} ions were introduced, AgNPs in the suspension became sparser and smaller (**Fig. 2A**, right), as more apparently shown in the amplified view of one particle (Insert, right top). Also, the hydrodynamic diameters of Gel-AgNPs before and after Hg^{2+} treatment were comparably examined by dynamic light scattering (**Fig. S2**, **ESI**), showing the average sizes decreased from about 13.0 nm to 8.5 nm. Furthermore, UV-vis spectra were also recorded for Gel-AgNPs with and without Hg^{2+} ions (**Fig. 2B** (a)). It was observed that Hg^{2+} ions could increase the UV-vis absorbance of Gel-AgNPs; yet, they could

make the yellowish Gel-AgNPs to fade out, as evidenced from their photographs (Insert, **Fig. 2B** (b) and (c)). The above results indicate that the mercury (II) etching of AgNPs in Gel-AgNPs could occur when they were treated with Hg^{2+} ions.

Furthermore, the enhanced catalysis mechanism for Hg^{2+} -stimulated catalysis activities of Gel-AgNPs was verified by essential experiments to make sure whether mercury (II) could etch Gel-AgNPs. Herein, Gel-AgNPs were deliberately treated UV lamp so as to grow AgNPs till their yellowish color changed to deep yellow. **Fig. S3A** (**ESI**) shows TEM images for the UV-treated Gel-AgNPs before and after the Hg^{2+} addition. A size-reducing change of AgNPs was observed for the latter. It could also be verified by the UV-vis spectra showing that UV-treated Gel-AgNPs could decrease in size after the Hg^{2+} etching, as shown in corresponding photographs (Insert) (**Fig. S3B**). On the basis of the above evidences, the mercury (II)-stimulated enhancement of peroxidase-like catalysis activities of Gel-AgNPs should result mainly from the decreased size of AgNPs by Hg^{2+} etching, of which the size-reducing AgNPs might achieve much higher catalysis activity, known as the “size effect”, as also evidenced elsewhere for gold and Fe_3O_4 nanomaterials³. Also, the special interaction between Hg^{2+} and AgNPs and the formation of Ag-Hg alloys changing the surface properties of AgNPs might be involved⁶, which remains to be further investigated in the future.

Moreover, the mercury (II)-enhanced catalysis dynamics of Gel-AgNPs was investigated in absence and presence of mercury (II) (**Fig. S4**, **ESI**). Here, colorimetric measurements were performed alternatively for different concentrations of TMB and H_2O_2 to obtain the Michaelis-Menten curves⁹ (**Fig. S4B** and **D**). It was found that Hg^{2+} -stimulated Gel-AgNPs presented much lower Michaelis constant (K_m) than Gel-AgNPs alone for TMB and H_2O_2 , together with larger maximal reaction velocity (V_{max}). The data indicate that Hg^{2+} -stimulated Gel-AgNPs could possess stronger catalysis activity especially higher affinity to TMB and H_2O_2 . Colorimetric analysis for Hg^{2+} ions could thereby be expected by Hg^{2+} -enhanced catalysis of Gel-AgNPs.

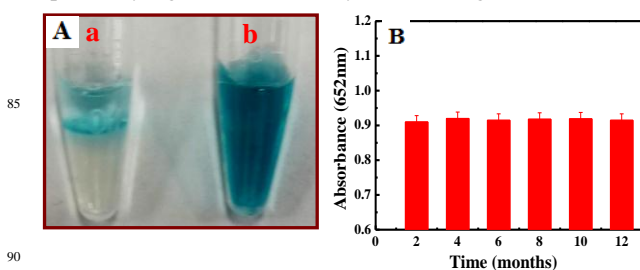


Fig. 3 (A) Comparable photographs of catalysis activities of Gel-AgNPs in TMB- H_2O_2 reactions with Gel-AgNPs originally (a) stored at $4\text{ }^\circ\text{C}$, and (b) thawed at $37\text{ }^\circ\text{C}$ with Hg^{2+} ions; (B) Mercury(II)-stimulated catalysis stability of Gel-AgNPs stored at $4\text{ }^\circ\text{C}$ in fridge over different time.

Moreover, the unique property of sol-gel transition of gelatin matrix (critical temperature at $35\text{ }^\circ\text{C}$) was investigated in switching the catalysis of Gel-AgNPs (**Fig. 3**). It was noted that Gel-AgNPs in gel phase (at $4\text{ }^\circ\text{C}$) with Hg^{2+} ions could not catalyze the TMB- H_2O_2 reactions, where the catalytic reactions occurred only on the gel surface (**Fig. 3A** (a)). When Gel-AgNPs were thawed at $37\text{ }^\circ\text{C}$ and diluted to be further mixed with Hg^{2+} ions, they could exert strong catalysis for the colorimetric reaction (**Fig. 3A** (b)), with the Hg^{2+} -stimulation time of about 5.0 min (**Fig. S5**, **ESI**). Particularly, such a temperature-

switching catalysis of Gel-AgNPs offered by gelatin matrix could allow for high catalysis stability of Gel-AgNPs, retained up to one year without a significant change (Fig. 3B). Importantly, Gel-AgNPs could be stably immobilized onto 96-well plates for long-term storage at 4 °C to facilitate colorimetric Hg²⁺ assays in label-free and high-throughput way. An advantage over other mercury (II) sensing assays, including these based on mercury-stimulated catalysis activities reported elsewhere⁶, could thus be expected.

Moreover, the dosage of gelatin for the synthesis of Gel-AgNPs was optimized as 4.0 % (Fig. S6A, ESI). Furthermore, the reaction conditions for the colorimetric Hg²⁺ analysis were optimized as 7.2 μM Gel-AgNPs (Fig. S6B), ionic strengths of 5.0 mM KNO₃, color reaction time of 20 min, neutral solution (pH 6.0 - 8.0) at room temperature (10-25°C) (Fig. S7, ESI), with 0.25 mM TMB and 10 mM H₂O₂ (Fig. S4A and C, ESI).

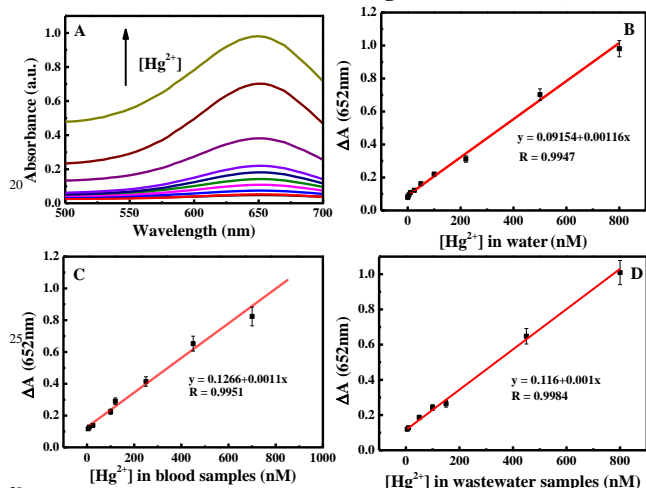


Fig. 4 Calibration detection curves for the linear relationship of absorbance values vs. different concentrations Hg²⁺ ions in samples of (B) water, (C) blood, and (D) wastewater, with typical UV-vis spectra of reaction products shown in (A). Colorimetric assays of Hg²⁺ ions in different samples were conducted by TMB-H₂O₂ reactions, corresponding to photographic results shown in Scheme 1.

Finally, calibration detection curves were obtained for Hg²⁺ ions with different concentrations in water, blood, and wastewater samples using 96-well plates (Fig. 4), corresponding to photographic results shown in Scheme 1. It was found that the typical UV-vis absorbance values could increase with increasing Hg²⁺ concentrations (Fig. 4A). A linear detection range of Hg²⁺ concentrations was obtained from 0.50 to 800 nM with the limit of detection (LOD) being 0.125 nM in water (Fig. 4B), which is lower than those of most other detection methods (Table S1, ESI). Moreover, the colorimetric detections of Hg²⁺ ions in blood and wastewater were conducted, with linear Hg²⁺ concentrations ranging from 5.0 nM to 700 nM and 1.0 nM to 800 nM, respectively. High detection selectivity for Hg²⁺ ions could also be achieved with no significant interference from other metal ions, as aforementioned in Fig. 1A. Therefore, the developed catalysis-based colorimetric assays could allow for the high-throughput Hg²⁺ detection with pretty high sensitivity and selectivity.

In summary, temperature-switchable gelatin matrix has been successfully employed for the one-pot synthesis of catalytic small AgNPs without any reductants. The resulting Gel-AgNPs could

display mercury-stimulated catalysis activity toward a simple, rapid, label-free, and high-throughput colorimetric protocol for probing mercury (II) in blood and wastewater using 96-cell plates. The catalysis-based detection mechanism involved was thought to result mainly from the mercury etching of AgNPs in gelatin matrix leading to smaller size of AgNPs with greatly enhanced catalysis activity for TMB-H₂O₂ reactions. The unique property of sol-gel transition of gelatin matrix could not only facilitate the temperature-switchable catalysis of Gel-AgNPs, but also allow for high catalysis stability of Gel-AgNPs. Remarkably, the catalysis-based method could detect Hg²⁺ ions in complicated wastewater and blood with high sensitivity, selectivity, and throughput. It might also circumvent some disadvantages of traditional detection approaches like fluorescence assays in terms of the detection stability and abilities against interferences from the complicated media. Such a colorimetric mercury assay promises huge potential applications in the clinical diagnosis, environmental monitoring, and pharmaceutical analysis fields.

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Notes and references

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