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ARTICLE

From Overlooked to Outstanding: A Molecular Silver Complex in Heterogeneous CO₂ Electroreduction

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Heterogenized molecular transition metal complexes, especially silver-based ones, have shown to be highly efficient as catalysts for electrochemical CO₂ reduction (eCO₂R). Herein we present the application of a silver dithiacyclam polymer (dithiacyclam = 1,8-dithia-4,11-diazacyclotetradecane, **Ag(dithiacyclam)**) in homogeneous as well as heterogeneous eCO₂R. Although the complex does not exhibit a noteworthy activity in solution, changing the catalyst environment through heterogenization onto a gas diffusion electrode (GDE) and subsequent application in a zero-gap electrolyzer (ZGE) drastically boosts its catalytic performance. At a current density of 50 mA/cm² a remarkable FE_{CO} of 97% is reached. Moreover, we highlight how optimization of the GDE fabrication *via* ink engineering including the choice of dispersion solvent results in a FE_{CO} up to 90% at an elevated current density of 300 mA/cm². Even at more application orientated current densities of 500 mA/cm² the eCO₂R outcompetes the competing hydrogen evolution reaction, achieving a FE_{CO} of 55%. Although signs of catalyst transformation into silver particles is observed in post-mortem analysis, these particles show higher activity than commercially available silver nanoparticles, thus highlighting that molecular systems can be very promising catalyst precursors for efficient eCO₂R.

Introduction

The electrochemical CO₂ reduction (eCO₂R) is a promising approach to sustainably convert the ubiquitous waste product CO₂ into valuable building blocks such as CO, ethylene and methane for the chemical industry.^{1,2} Depending on the investigation goals – fundamental understanding or application orientated performance – the choice of electrolyzer type plays a crucial role. Batch- or H-Type electrolyzer are commonly applied to test intrinsic catalyst properties, either in a dissolved or heterogenized form. However, these systems are not suitable for scale up as they are not designed for continuous-flow operation. In addition, they exhibit several limitations such as high cell resistances caused by the liquid electrolyte, resulting in increased energy consumption, and CO₂ mass transport limitations arising from its limited solubility in the chosen electrolyte. A scalable zero-gap electrolyzer (ZGE) omits the previously mentioned disadvantages by operation without liquid electrolyte compartments.^{3–6} Further it has been shown to be able to operate under industrially relevant conditions of high current densities (>200 mA/cm²), elevated temperatures (>60 °C) and cell voltages < 3 V.^{1,7,8} Commonly a gas diffusion electrode (GDE), is employed as cathode, in which the catalyst is coated onto a layer of woven carbon fibres, the so called gas diffusion layer (GDL), which is directly supplied with gaseous CO₂. The catalyst layer (CL) commonly consists of the catalyst, a conductive carbon support and an

ionomeric binder. This GDE is directly in contact with the anode *via* an ion-permeable solid polymer membrane, which is commonly an anion exchange membrane (AEM) (Figure 1).^{2,6,9–11}

Previous studies over several decades have demonstrated that transition metal complexes exhibit a high activity for eCO₂R in homogeneous catalysis.^{12–15} As the research on heterogenized molecular catalysts emerged during recent years, it was found that the high activity achieved in solution is not necessarily transferable onto other electrolyzer types like a ZGE. This behaviour was recently highlighted for iron tetraphenyl porphyrin (FeTPP) which is well known for its homogeneous eCO₂R activity.^{16–21} Yet, it collapsed once implemented into a ZGE even at low current densities ≤100 mA/cm².^{20–22} Contrarily, molecular silver complexes which are not known to homogeneously catalyse the eCO₂R achieved remarkable performances applied in ZGEs.^{6,12,22,23} A good example to showcase the efficiency of silver complexes in ZGEs are i.e. silver-based BIAN complexes (BIAN = bis(imino)acenaphthene) which reached a remarkable faradaic efficiency for CO (FE_{CO}) of 67% at a current density of 600 mA/cm² after careful optimization of the GDE fabrication.¹⁰

For the numerous examples of molecular electrocatalysts in eCO₂R it is well established that the performance is highly dependent on its environment: in homogeneous electrolysis this mainly comprises chosen solvent, catalyst concentration and kind and concentration of proton source.^{12,21} Whereas immobilized complexes in ZGEs are affected by a variety of different parameters, such as catalyst loading, nature and amount of carbon support, the dispersion medium, or the binder composition, which determine the local catalyst environment in the CL. Besides, process parameters like

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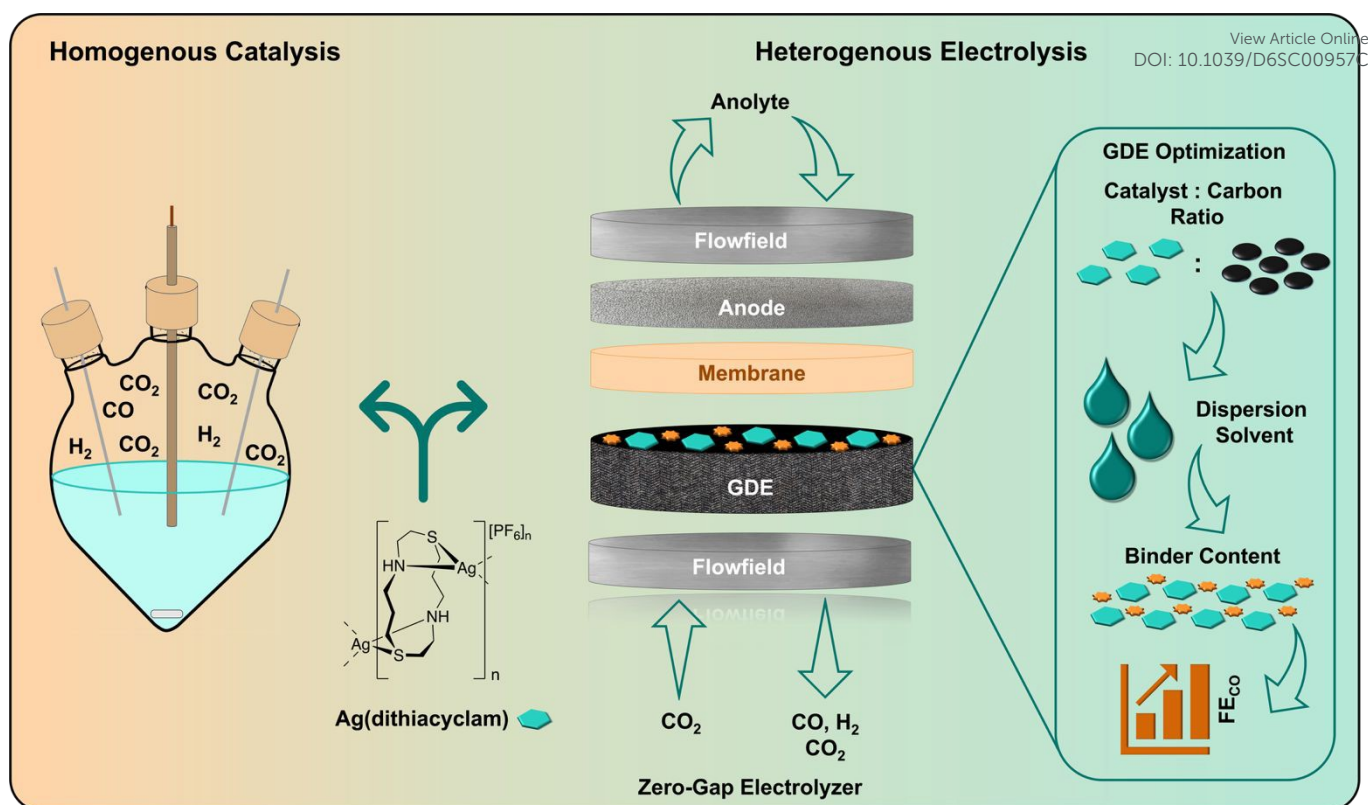


Figure 1. Graphical overview of the herein presented work: tuning the electrolysis environment of **Ag(dithiacyclam)** from a homogeneous solution to a heterogeneous environment within a GDE along with tailoring of the GDE composition to maximise the eCO₂R activity.

applied current density, anolyte composition, gas flow or humidification thereof influence the catalysis and performance.^{6,10,24,25}

While individual studies have addressed isolated some of the mentioned parameters, mainly in metallic-systems, a holistic optimization approach under industrially relevant conditions ($\geq 200 \text{ mA/cm}^2$, $\geq 60 \text{ }^\circ\text{C}$) for molecular-ones remains unexplored.^{26,27} Hence, we sought to present a comprehensive study of a previously untested molecular catalyst for eCO₂R encompassing both its homogeneous catalytic activity and its performance under application-relevant ZGE conditions. To design this study, we took advantage of the extensive literature on molecular transition metal complexes and selected the previously described **Ag(dithiacyclam)** complex (dithiacyclam = 1,8-dithia-4,11-diazacyclotetradecane) motivated by three main considerations. First, the cyclam ligand system (1,4,8,11-tetraazacyclotetradecane) is well established, and complexes derived from it have been thoroughly investigated as homogeneous eCO₂R catalysts due to their high efficiencies in aqueous solutions.^{28,29} However, their use in ZGEs, particularly in combination with commonly applied AEMs, has been rarely explored.³⁰ This limited application is largely due to the high aqueous solubility of transition metal complexes of the standard cyclam ligand, which is unsuitable for ZGE catalysis because the catalyst may be washed out of the GDE.³¹ To overcome this limitation, we selected the sulfur-containing cyclam analogue dithiacyclam, as its complexes are typically water-insoluble while remaining active in eCO₂R.^{32,33} At last, we chose silver as the metal centre because silver-based molecular systems are currently the best-performing catalysts in ZGEs, and the coordination environment of dithiacyclam is well

suited to accommodate the soft Ag(I) ion.^{22,24,34} As reported previously, dithiacyclam forms water-insoluble polymeric **Ag(dithiacyclam)** structures upon coordination with Ag(I), making this complex particularly well suited for the purpose of our study.

Thus, we herein present a study comprising homogeneous and heterogeneous investigations using **Ag(dithiacyclam)** as catalyst in eCO₂R. Our findings underscore not only the importance of probing catalytic systems directly in heterogeneous environments but also how much potential performance may be overlooked when relying solely on homogeneous catalyst screening. Exploring the difference of homogeneous and heterogeneous operation opens a largely untapped field with significant implications for catalyst development and industrial electrochemical processes.

Results and discussion

Homogeneous Electrochemistry

The herein investigated **Ag(dithiacyclam)** was prepared according to a literature known procedure.³⁴ Electrochemical investigations thereof were performed in a one compartment cell using a glassy carbon (3 mm diameter) working electrode, a platinum wire as counter electrode and a silver wire as pseudo reference electrode. Ferrocene (Fc) was added as internal reference and all cyclic voltammograms (CVs) were referenced against the Fc/Fc⁺ redox couple. CVs of the complex were recorded in a 1 mM acetonitrile (MeCN) or MeCN:H₂O (4:1) mixture either in an inert Ar atmosphere or a CO₂ atmosphere using 0.1 M tetrabutylammonium hexafluorophosphate (TBAPF₆) as supporting electrolyte and show a similar



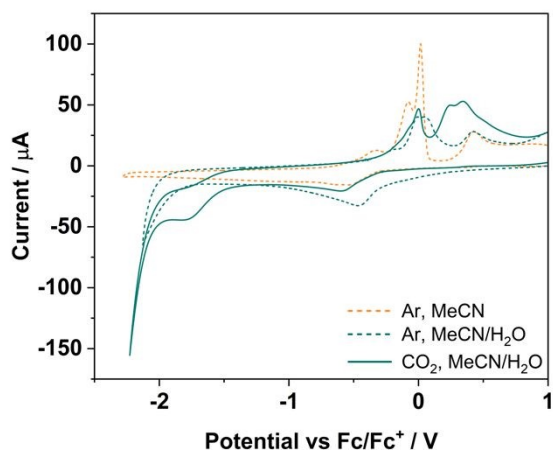


Figure 2. CVs of **Ag(dithiacyclam)** (0.1 mM) recorded at a scan speed of 100 mV/s with 0.1 M TBAPF₆ as supporting electrolyte in pure MeCN in an Ar atmosphere (orange dashed line) or in MeCN/H₂O 4:1 with 0.1 M TBAPF₆ as supporting electrolyte, either in an Ar (dark green dashed line) or CO₂ (dark green line) atmosphere.

electrochemical behaviour as the non-sulphur containing cyclam variant.^{31,35} (Figure 2, Figure S1) In pure MeCN one reductive wave at -0.55 V vs Fc/Fc⁺ is observed which can be attributed to the Ag(I) to Ag(0) reduction. This reduction is not completely reversible, the oxidative wave at -0.33 V vanishes with higher scan rates (Figure S2). Around 0 V vs Fc/Fc⁺ a strong stripping signal is observed which can be attributed to the previous formation of a Ag(0) deposit on the surface of the working electrode which is reoxidized to Ag(I) and Ag(II). The irreversible oxidation observed at +0.42 V likely corresponds to the Ag(I) to Ag(II) transition, whereas the metal ion is likely to leave the coordination environment as rather soft coordination environment of dithiacyclam is not suitable anymore for a hard cation. Purging the complex solution in pure MeCN with CO₂ does not alter the CV. The addition of H₂O in an Ar atmosphere shifts the Ag(I) reduction to a potential of -0.48 V, additionally a beginning catalytic current, attributed to proton reduction, is observed at a potential of -2.11 V vs Fc/Fc⁺. In presence of CO₂, a new reductive wave, induced by interaction with CO₂, evolves at -1.75 V vs Fc/Fc⁺ followed by a catalytic current. Additionally, a newly formed oxidative wave at +0.25 V vs Fc/Fc⁺ can likely be attributed to the oxidation of a silver-(bi)carbonate species, as this signal is only observable in presence of protons and CO₂, as also observed for other silver complexes.²³

The ability of **Ag(dithiacyclam)** to reduce CO₂ in solution was investigated *via* controlled potential coulometry (CPC) at the potential of the second reductive wave in presence of water and CO₂, -1.85 V vs Fc/Fc⁺. (Figure S3) Catalysis was then performed for a duration of 24 h. The gaseous headspace was analysed two-hourly for the first 8 h and after 24 h *via* gas chromatography (GC). Only negligible quantities of CO with a FE_{CO} of 2%, were observed after 24 h with large variations and poor reproducibility. The drop in total FE after 8h to around 70% can be explained by the formation of (bi)carbonates, as known for silver in acetonitrile, as well as electrolyte and catalyst degradation (*vide infra*).^{36,37} After electrolysis, a deposit was

visible on the surface of the working electrode, which was verified to be elemental silver *via* scanning electron microscopy with coupled energy dispersive X-ray spectroscopy (SEM/EDX) (Figure S4). As CO is solely evolving after more than 8 h of electrolysis, the formation of CO is most likely generated by this silver deposit, as also verified by a performed rinse test (Figure S5). An electrode deposit was already observed after 2 h of electrolysis, as checked *via* SEM/EDX this species is likely deposited **Ag(dithiacyclam)** as the EDX analysis revealed a clear overlap of silver, nitrogen and sulphur (Figure S6), whereas the ligand breaks down in the ongoing electrolysis, indicating that the observed CO formation can be attributed to the elemental silver deposit formed later on. Thus, the rapid deposition accompanied by decomposition of the silver complex might have caused the big variations in FE as similar behaviour was previously observed for other molecular catalysts in solution.³⁸

Heterogeneous Electrochemistry

Traditionally, the lack of activity of **Ag(dithiacyclam)** in homogeneous catalysis would discourage further investigation. This paradigm stems from the common practice of screening molecular complexes in homogeneous systems before translating them to heterogeneous eCO₂R, thereby overlooking candidates that may only become active upon immobilization.^{6,39} Nevertheless, we pursued further studies on heterogeneous catalysis in ZGEs as we think **Ag(dithiacyclam)** is a promising complex for implementation into ZGEs. This assumption is supported by other molecular silver complexes that show high activity in ZGEs despite being only weakly active in solution, suggesting that **Ag(dithiacyclam)** may exhibit a similar behaviour.^{10,22,24}

Electrolysis at low current densities (<100 mA/cm²). First electrolysis experiments were performed in a ZGE equipped with a PiperION AEM (40 μm, Versogen, conditioned in 1 M KOH), a Ni-foam anode and 1 M KOH cycled as anolyte (20 mL/min). The cathode was fed with humidified CO₂ (20 mL/min, 100% rel. humidity) and the outlet gas stream was directly connected to a gas chromatograph (GC). Subsequent to a short electrochemical conditioning of the electrodes, (details in SI) electrolysis at 50 mA/cm² was conducted for 1 h. The product gas composition was analysed every 30 min of electrolysis *via* inline GC. GDEs have been prepared by drop casting a catalytic ink onto a carbon cloth, the detailed procedure is described in the SI. In a first trial a **Ag(dithiacyclam)** loading of 0.5 mg/cm² accompanied by 1.0 mg/cm² carbon black (ENSACO 250G) as carbon support and 0.2 mg/cm² SustainION XA-9 binder were chosen as GDE composition based on previous reports (Figure S7).²⁴ In strong contrast to the almost absent activity of the catalyst in solution, a high activity in eCO₂R was observed reaching a remarkable FE_{CO} of 88% at 50 mA/cm². *Via* changing the catalyst environment from an organic solvent to a solid, hydrophobic carbon matrix, the eCO₂R behaviour of **Ag(dithiacyclam)** is completely changed. This activity increase can be attributed to



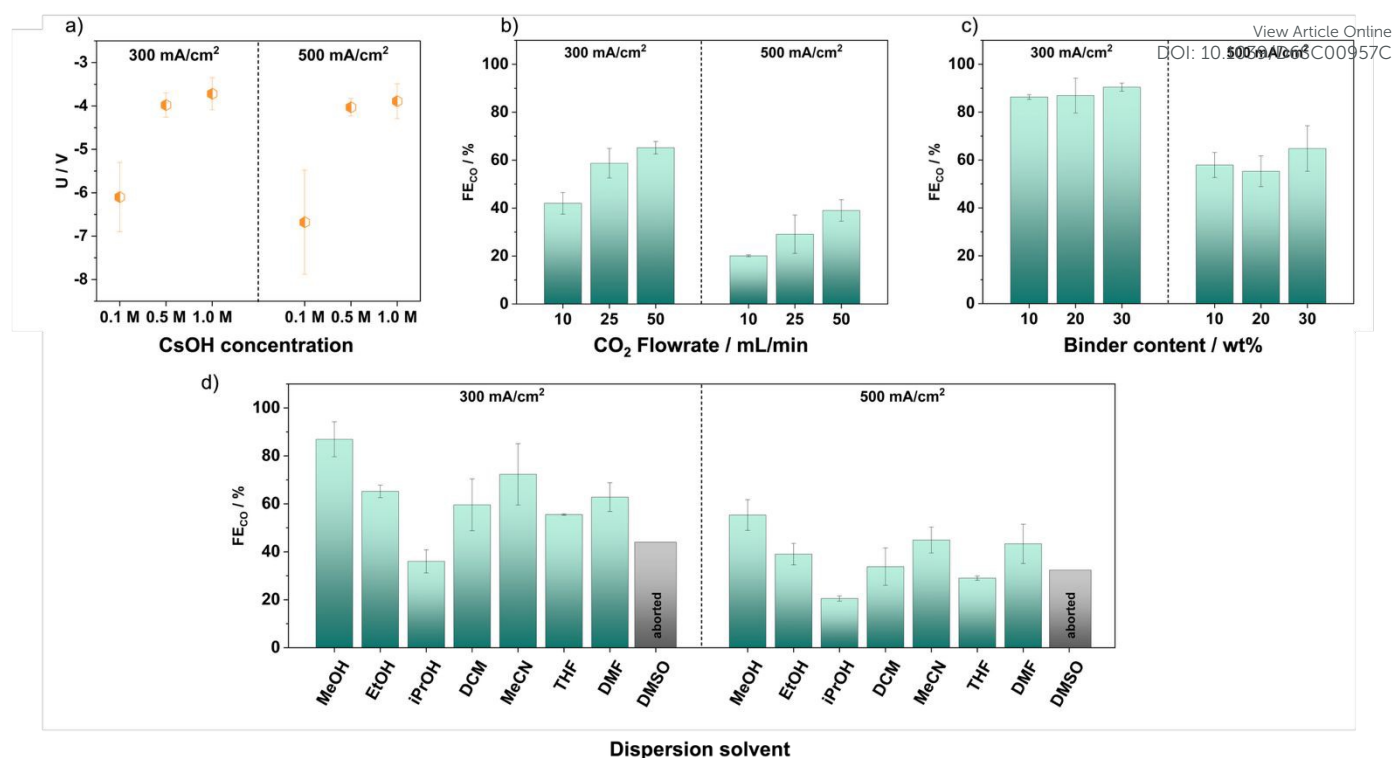


Figure 3. Selection of outstanding measurement data from electrolysis optimization processes at 60°C at current densities of 300 mA/cm² and 500 mA/cm² achieved with a GDE comprising a **Ag(dithiacyclam)** loading of 0.5 mg/cm². a) Averaged cell voltages U achieved while applying different CsOH concentrations; FE_{CO} achieved while b) applying the given λ_{CO_2} values; c) comprising the given binder content in the CL; d) utilizing the given solvent for dispersion of the catalytic ink.

beneficial properties of embedding molecular catalysts into GDEs: a very high local CO₂ concentration, improved electron transfer by the carbon matrix and suppression of the competing HER by the alkaline and rather hydrophobic environment.^{40–42} Such a discrepancy in catalytic eCO₂R activity between homogeneous electrolysis and zero-gap catalysis is rarely observed, since most molecular catalysts reported to be active in zero-gap electrolyzers, such as CoTPP or Ag(BIAN)s, also display at least moderate activity in homogeneous electrolysis.^{22–24,43}

Since we want to exploit the full potential of **Ag(dithiacyclam)** as eCO₂R catalyst in GDEs, we wanted to determine the ideal carbon support as well as the best catalyst:carbon ratio since no clear trend exist in literature – another issue of the minimal amount of holistic reports.⁶ Hence, we tested two different carbon black variants (ENSACO 250G and SuperP) which alter in key characteristics such as porosity, electrical conductivity and dispersibility in the solvent of catalytic inks.^{10,44–46} Besides carbon black we investigated multi walled carbon nanotubes (MWCNTs) as they also exhibit a good conductivity.^{6,22} Under laboratory scale conditions as applied here, all investigated carbon supports and ratios thereof exhibit equally good performance reaching a FE_{CO} around 90% (Figure S7). Nevertheless, a trend towards higher FE_{CO} is observed when SuperP with a loading of 0.5 mg/cm² is applied as carbon support reaching a maximal value of 97% which can be attributed to the high conductivity and large pore size.^{44,45} Thus, SuperP was chosen for following studies. At lower current

densities, the intrinsic activity of **Ag(dithiacyclam)** is sufficiently high that parameter-specific effects become difficult to disentangle, limiting the ability to draw meaningful conclusions. Therefore, the influence of varying GDE and process parameters was evaluated at higher current densities.

Minimizing the catalyst loading (≥ 100 mA/cm²). After proving to be a suitable catalyst for eCO₂R in ZGEs at low current densities (≤ 100 mA/cm²) further experiments targeting more industrially relevant current densities up to 500 mA/cm² were performed.¹ To reduce the chance of carbonate formation on the cathode side, which can block the pores of GDEs, the anolyte was changed to 1 M CsOH cycled with 20 mL/min.^{47,48} Additionally, the anode was exchanged to an IrO₂ coated Ti-felt (1 mg/cm²) due to its high activity as oxygen evolution catalyst.^{49,50} Molecular catalysts offer one big advantage compared to pure metal particles: they reach equal or even higher FE_{CO} values while bearing an overall lower metal content in the CL and thus reduce the need of precious metals.²⁴ At room temperature three different catalyst loadings (0.05 mg/cm², 0.125 mg/cm² and 0.25 mg/cm²) have been tested while keeping the mass of SuperP carbon black (0.5 mg/cm²) and Binder (SustainION XA-9, 0.4 mg) constant (Figure S8). After an electrochemical conditioning of the electrodes consecutive electrolysis was performed at 100 mA/cm², 300 mA/cm² and 500 mA/cm² for 30 min each following a product gas analysis *via* inline GC. At 100 mA/cm² no significant difference in FE_{CO} is observable and selectivity close to 100% CO is generated for all loadings. Nevertheless, the performance rapidly decreased with low catalyst loadings as the



current density was increased. The GDE with a loading of 0.25 mg/cm² achieved the highest FE_{CO} of 28% at 500 mA/cm² and the GDE with 0.125 mg/cm² reached the lowest FE_{CO} of 9%. Yet, the measured cell voltage also increases with a higher catalyst loading, reaching -5.25 V. As a control experiment GDEs with Ag NPs (< 40nm) with an equal metal content have been prepared and tested under identical conditions leading to smaller FE_{CO} at elevated current densities of 300 and 500 mA/cm² (Figure S9) yet maintaining a lower cell voltage than **Ag(dithiacyclam)**. Since the ligand scaffold ensures a heteroatom rich environment, which herein benefits the CO generation activity, further testing towards even more industrially relevant conditions is pursued to assess transferability, identify potential mass-transport limitations, and evaluate catalyst stability.

Process optimization for electrolysis at 60 °C. As industrial scale electrolysis beneficially to be run at elevated temperature, we performed the following catalysis experiments at 60 °C and high current densities of 300 mA/cm² and 500 mA/cm².¹ These experiments were performed with a catalyst loading of 0.5 mg/cm². The SuperP and SustainION XA-9 loadings of 0.5 mg/cm² and 0.2 mg/cm², respectively, were maintained for process optimization studies. A humidified (80% rel. humidity) CO₂ flow of 50 mL/min was fed to the cathode as starting point. At first the influence of cation crossover through the membrane from the anolyte was analysed by testing CsOH in concentrations of 0.1 M, 0.5 M and 1 M CsOH (Figure S10). In terms of FE_{CO} all anolytes achieved values of 60%±7%, however, a decrease in cell voltage is observed with low concentrations leading to averaged cell voltages from -6.0 V, -3.85 V and -3.75 V at 300 mA/cm² respective to concentrations of 0.1 M, 0.5 M and 1.0 M caused by lower solution resistance and preferred conditions for the oxygen evolution reaction (Figure 3a).^{25,51,52} At 500 mA/cm² the differences are even more prominent reaching a cell voltage of -6.5 V for 0.1 M CsOH and -3.9 V for 1.0 M CsOH (Figure S10).

To ensure efficient proton supply, ionic conductivity and stabilization of the gas-liquid-solid interface required for efficient eCO₂R, the next step of process optimization focused on the water delivery towards the cathode. Thus, different humidification percentages (0%, 25%, 50%, 80% and 100% rel. humidity) have been investigated by controlled variation of the bubbler temperature (Figure S11). Below a humidification degree of 80% a FE_{CO} of 50% at 300 mA/cm² is not exceeded indicating an insufficient water delivery. At 80% rel. humidity a maximum FE_{CO} of 65% is achieved. With an even higher humidification of 100% the FE_{CO} drops to 50% accompanied by a high cell voltage of -8.5 V most likely caused by electrode flooding due to large water amounts.

Aiming towards industrially relevant conditions, an efficient CO₂ utilization is desired. Thus the λ_{CO2} value – the ratio between supplied and actually converted CO₂ – should be as low as possible, yet it remains an uncommon metric in studies of molecular catalysts.^{53,54} Therefore, we decreased the CO₂ feed to the cathode from 50 mL/min to 25 mL/min and 10 mL/min (Figure 3b, Figure S12). The corresponding λ_{CO2} values at the applied current densities can be found in Table S1. With decreasing λ_{CO2} values the

selectivity for CO generation decreases, reaching only 20% and 29% FE_{CO} as observed for other molecular silver complexes in GDEs.⁷⁴ When the cathode feed stream is set to 10 mL/min, the CO₂ delivery to the cathode is insufficient as indicated by a FE_{total} of solely 73%. Even though high λ_{CO2} values of 12 and 7 are needed to achieve good FE_{CO} values of 65% and 39%, respectively, it needs to be considered that the system is operating with a low silver loading of 0.22 mg/cm² which is much lower than the commonly reported values of several milligrams per square centimetre of silver nanoparticle (NP) based GDEs.^{54–57} This low loading likely requires high CO₂ feed rates to ensure that the generally low number of active sites is constantly supplied with CO₂, to avoid local CO₂ depletion and concentration gradients at the catalyst surface.

Optimization of dispersion solvent. After determining the optimal process parameters for catalysis at 60 °C, the GDE was optimized *via* ink engineering to control the electrode structure for optimal ion and electron transport and overall performance. Thus, the dispersion solvent and binder content have been varied and correlated with microstructure analysis of the GDEs. For all samples, 5 mg Ag(dithiacyclam) and 5 mg SuperP were dissolved in 2 mL solvent as before. Besides the primarily utilized EtOH, seven solvents of varying polarity were investigated (MeOH, iPrOH, MeCN, THF, DMF, DCM, DMSO) while keeping the binder content constant at 20wt% (Figure 3d, Figure S13). These solvents were chosen based on previous reports of dispersed molecular catalysts as well as good solubility (DCM, DMSO, MeCN) of **Ag(dithiacyclam)**.^{10,43,58} When DMSO was used, the measurements were aborted due to cell voltages exceeding 10 V during the first 20 min of electrolysis (Figure S13h). While disassembling the cell it was observed that the GDE surface was “glued” to the membrane, potentially caused by interaction of leftover DMSO with the AEM, as the electrodes were only dried at 95 °C to maintain the integrity of SustainION XA-9 (Figure S14). Thus, DMSO was excluded as suitable solvent for scale up processes. At 300 mA/cm² the lowest FE_{CO} was achieved utilizing iPrOH with only 32%. As the observed variations in FE_{CO} are likely attributed to structural changes in the catalyst layer of the GDE, we performed a combination of SEM/EDX and computer tomography (CT) for microstructure analysis. As observed in both methods the CL of the iPrOH GDE is rather loosely packed (Figure 4c) and comprises several big particles agglomerates (Figure S15-16). Especially the latter can explain the comparably bad performance in eCO₂R as not many catalytic centres are accessible for the CO₂ (Figure S16) All other tested solvents, except MeOH, range between FE_{CO} values of 56% and 72%. Of those mediocre performing GDEs, we decided to analyse the EtOH-based one *via* SEM/EDX and CT, as a FE_{CO} trend (higher FE_{CO} with higher polarity) is observed within the protic solvents. The CL of the EtOH-based GDE is packed a little denser (Figure 4b), which can be beneficial for the conductivity (Figure S17-19). In the as-prepared state, fewer particle agglomerates are observed compared to the GDE prepared with iPrOH (Figure S19). However, post-mortem analysis reveals that a larger number of particle agglomerates form in the EtOH



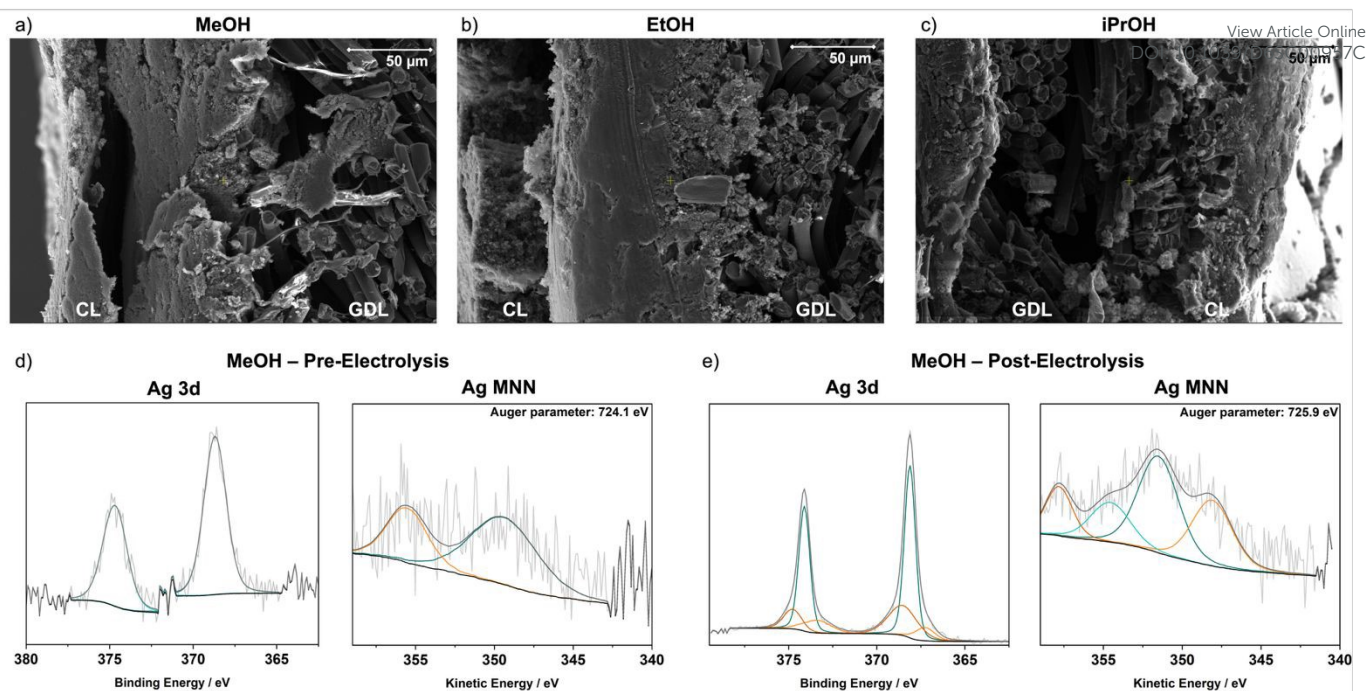


Figure 4. Recorded cross sectional SEM images at a magnification of 800x of pristine **Ag(dithiacyclam)** GDEs prepared with a) MeOH; b) EtOH; c) iPrOH. Recorded Ag 3d and Ag NMM XPS spectra of d) a pristine GDE prepared with MeOH and e) a utilized GDE prepared with MeOH applying in short term catalysis at 300 mA/cm² and 500 mA/cm².

based GDE during electrolysis (Figure S20-21). From all studied dispersion solvents MeOH achieved the highest FE_{CO} of 87% at 300 mA/cm² and even at a current density of 500 mA/cm² the eCO₂R dominates reaching a FE_{CO} of 55%. These findings are in line with previous reports, which determined MeOH as most suitable dispersion solvent for a polar silver complex.¹⁰ The transferability of dispersion properties from one complex towards another simply based on the complex' polarity is a crucial finding to efficiently fine tune GDE fabrication. SEM/EDX and CT analysis of a MeOH-based GDE revealed that this high activity is likely caused by a very thin (Figure 4a) and densely packed CL without visible agglomerates (Figure S22-24). Thus, a good CO₂ distribution as well as conductivity are ensured. After electrolysis still no silver particles were observed maintaining an even CL surface (Figure S25-26). However, it needs to be noted that for the latter a big part of the CL was removed, most likely during sample preparation, as this is also clearly observed in the CT images.

Besides the structural electrode features, the complex' integrity is of equal if not even greater interest. To analyse this, we decided for XPS spectroscopy, as we can determine the metal's oxidation state and coordination environment (Figure 4d-e, Figure S27-30). No matter if the GDE was prepared with MeOH or EtOH the pristine GDEs show Ag 3d_{5/2} signals at 368.7/368.8 eV with Auger parameters of 724.1 and 723.8 eV, which is in line with XPS spectra reported for Ag(I) complexes on GDEs before.^{10,24} In addition the spectrum of the S 2p orbital shows signals for C-S-C as well as signals that can be attributed to a S-Ag bond.⁵⁹ Post-electrolysis GDEs revealed the decomposition of the initial complex for both preparation methods, as three silver species are observed in the Ag 3d_{5/2}

orbital spectra which can be attributed to a mixture of **Ag(dithiacyclam)**, AgO and elemental silver (Figure S29-30).^{60,61} In contrast to the degradation process in homogeneous electrolysis, leading to an elemental silver surface, the *in situ* generated silver particles within the GDE are supported by a heteroatom rich carbon matrix, which can enhance eCO₂R activity.^{62,63} Further, we would like to note that the observed silver species are not necessarily the species catalysing the eCO₂R, as they potentially might be reoxidizing after the electrolysis is stopped, as also observed for other molecular catalysts.^{64,65} Even though a transformation of **Ag(dithiacyclam)** is observed, the silver species formed exhibits an elevated point towards eCO₂R. Thus, the application of silver coordination compounds as precatalysts for eCO₂R in ZGDEs could become a promising approach to lower the silver content in GDEs. As we have recently highlighted, the final species – whether molecular or precatalyst – has little importance on the industrial stage, as long as the necessary target values are achieved.^{6,10}

Optimization of binder content. Lastly, the influence of binder content, previously equal to 20 wt%, was in- and decreased to 10 wt% and 30 wt% (Figure 3c, Figure S31). With a binder content of 30 wt%, a slight increase of the FE_{CO} to 90% at 300 mA/cm² was observed, however at 500 mA/cm² a noteworthy increase to 65% was achieved. This improvement is likely attributed to a very dense yet highly porous surface structure combining excellent CO₂ transport with high conductivity (Figure S32). In contrast, no significant changes in FE_{CO} as well as electrode structure were observed with the lower binder content of 10 wt% (Figure S33). Therefore, all subsequent experiments were conducted using 30 wt% binder, as this composition proved to be the most effective. Under the herein optimized conditions, GDEs prepared with silver NPs and



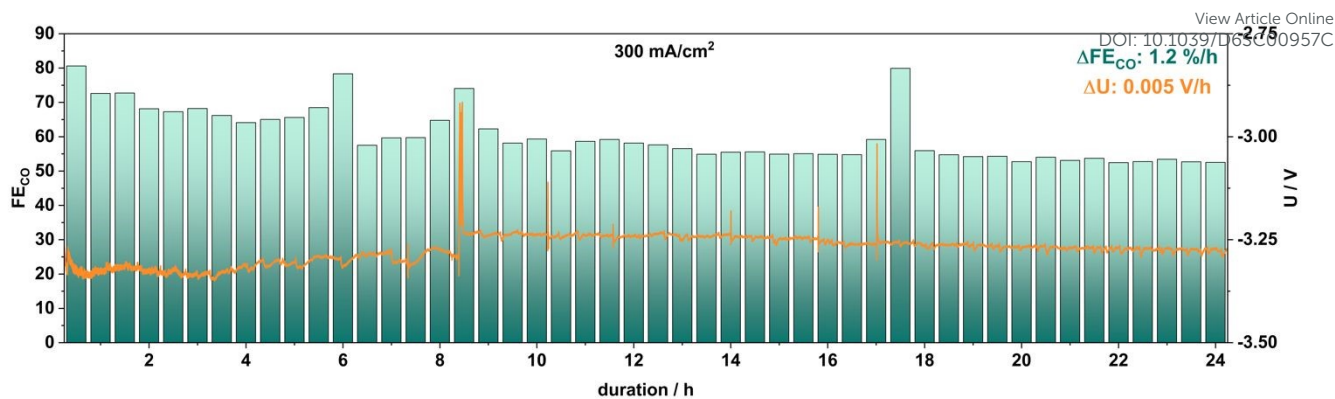


Figure 5. Measured FE_{CO} (green bars) and cell voltage U (orange line) over an electrolysis duration of 24 h applying an optimized GDE with a **Ag(dithiacyclam)** loading of 0.5 mg/cm^2 and a binder loading of 30 wt% which was prepared using MeOH as dispersion solvent. The decay rates of FE_{CO} and U are given in the graph.

carbon additive maintaining an equal silver loading to that of **Ag(dithiacyclam)** did not reach the FE_{CO} achieved by the molecular system reaching only FE_{CO} values of 26% and 9% at 300 mA/cm^2 and 500 mA/cm^2 , respectively. (Figure S34). Overall, this stepwise optimization of electrolysis parameters and the GDE composition establishes **Ag(dithiacyclam)** as one of the most efficient molecular systems reported in ZGEs, while also enabling performance comparable to that of heterogeneous silver-based materials at high current densities (Table S2).^{22,24,66–72}

Long-term catalysis. Currently there is still a big gap in research how molecular catalysts perform over several hours in ZGEs at industrially relevant current densities, as most studies have been performed at current densities $\leq 200 \text{ mA/cm}^2$ and room temperature.⁶

Hence, we tested the optimized GDEs over a time span of 24 h at 300 mA/cm^2 (Figure 5, Figure S35). Even after 24 h of constant electrolysis the FE_{CO} was still above 50%. Within the first 8 h of electrolysis the cell voltage fluctuates between -3.35 V and -3.3 V , following a quick jump towards -3.25 V . Since this jump is accompanied by a total FE exceeding 100% by far it can be assumed that this was caused by a pressure drop i.e. caused by temporary blocking of tubes by water droplets or carbonate salts. A similar jump is again observed after 16 h. Nevertheless, the cell voltage only slowly decreases from -3.23 V (8h) to -3.27 V after 24 h, corresponding to a cell voltage decay rate of -0.005 V/h .

Surface analysis of the GDE applied in 24 h electrolysis *via* SEM/EDX showed small, isolated silver agglomerates in elemental mapping of the EDX without any traces of ligand (Figure S36). CT analysis revealed that the GDL was blocked by particles, probably carbonate salts and CsOH (Figure S37). This GDE flooding might have caused the drop in FE_{CO} as CO_2 cannot be delivered sufficiently to the CL. Future efforts to improve long-term stability can focus on the reduction of salt formation, i.e. *via* utilization of a bipolar membrane in reverse-bias configuration or reduction of the cation concentration within the anolyte.⁷³ Similar to the GDEs utilized in short-term testing, XPS analysis of the Ag 3d orbital revealed a mixture of **Ag(dithiacyclam)**, AgO and elemental silver after 24 h of electrolysis (Figure S38). Noteworthy, the formed mixture of

silver compounds exhibits a higher activity after 24 h than a pristine electrode prepared with Ag NPs and carbon at the same metal loading, which solely reaches an FE_{CO} of 26% after 30 min of electrolysis (Figure S34).

Overall, these promising values show how the direct incorporation of **Ag(dithiacyclam)** into GDEs gives another example how efficiently transition metal complexes can perform in ZGEs under industrially relevant conditions, urging the community to further understand its importance as the first step to testing molecular electrocatalysts. Further, we have shown that a usually overlooked component, the dispersion solvent, can drastically tune a catalysts performance as FE_{CO} values between 32% and 87% have been observed *via* changing the dispersion solvent as it influences the microstructure of the CL drastically. One final crucial milestone to achieve is long-term stability, which can be further improved by profoundly tailoring the binder type and alkali-ion concentration at the GDE interface, a topic that has only been investigated for metal nano-particles.^{41,74–76}

Conclusions

Herein we have shown that the surprising trend that a transition metal complex, which is not homogeneously catalysing the $e\text{CO}_2\text{R}$ can nevertheless be a very potent catalyst in heterogeneous $e\text{CO}_2\text{R}$. In essence the results redefine how we should test and optimize molecular electrocatalysts for potential large scale applications. Direct laboratory scale testing in a ZGE showed that **Ag(dithiacyclam)** can convert CO_2 to CO efficiently with a FE_{CO} of 97% at a current density of 50 mA/cm^2 . In first electrolysis experiments applying industrially relevant conditions (60°C , $\geq 300 \text{ mA/cm}^2$) a FE_{CO} of 65% was achieved at 300 mA/cm^2 . Notably, to maximise the activity of **Ag(dithiacyclam)** the catalytic ink composition was varied, as this is a crucial part of GDE fabrication which is often overlooked in the molecular community. Especially the choice of dispersion solvent impacts the $e\text{CO}_2\text{R}$ performance of a GDE drastically resulting in a broad range of FE_{CO} covering values between 32% and 87% for eight different solvents. Of those MeOH reached the highest FE_{CO} of 87% which was further improved to 90% *via* increasing the binder content.



With the optimized GDE composition, a remarkable FE_{CO} of 65% was achieved at 500 mA/cm² in short term testing and a constant FE_{CO} above 50% was maintained for 24 h at 300 mA/cm². Via SEM and CT analysis it was found that the dispersion solvent influences the structure of CL and the formation of catalyst aggregates within. XPS analysis revealed partial transition of the molecular structure into a mixture of highly active silver particles which are more active than commercial silver NPs.

Hence, the **Ag(dithiacyclam)** is a great example for the big potential of silver coordination compounds as catalyst precursor for eCO₂R in ZGEs. On a further notice our study points out important research questions which remain unanswered regarding the nature of the formed catalyst species and how they interact with the carbon support material. Moreover, this work stands to highlight how testing molecular electrocatalysts directly in heterogeneous conditions could be even more directed towards industrially relevant conditions. Particularly, placing the focus on catalytic ink optimization, as it drastically governs the CL structure and, as a result, the eCO₂R performance as well.

Author contributions

W.W.: conceptualization, formal analysis, investigation, methodology, validation, visualization, writing – original draft. K.P.: writing – review & editing. J.J.: supervision, writing – review & editing. U.-P.A.: conceptualization, supervision, funding acquisition, resources, writing – review & editing.

Conflicts of interest

There are no conflicts to declare.

Data availability

The data supporting this article have been included as part of the Supplementary information.

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- The data supporting this article have been included as part of the Supplementary Information. Supplementary information: Furthering electrochemical data, CT data, SEM/EDX data, XPS data, NMR spectra and further experimental details. See DOI: [URL – format <https://doi.org/DOI>]

