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Diketo[*n*]CPPs as chiral and shape-adaptive fullerene hosts and precursors to DBP[*n*]CPPs

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Conjugated nanohoops, such as [*n*]cycloparaphenylenes ([*n*]CPPs) and derivatives, exhibit unique structural and optoelectronic properties, making them promising candidates for applications in optoelectronic materials, and as hosts for supramolecular chemistry. Using π -systems unsymmetric to rotation or incorporating chiral units can furnish chiral nanohoops. We herein present the synthesis and characterization of diketo[8]- and diketo[9]CPPs, along with their corresponding dibenzo[*a,e*]pentalene (DBP) derivatives, DBP[8]- and DBP[9]CPP. Due to the central chirality of the diketone-units, these nanohoops are chiral without the possibility of racemization through rotation and show distinct chiroptical properties. The diketo[*n*]CPPs possess high fluorescence quantum yields of 87% ($n = 8$) and 92% ($n = 9$). The shape-adaptive properties of diketo[*n*]CPPs, facilitated by the tunable kink angle of the diketo unit, enable efficient host–guest interactions with fullerenes. Fluorescence titration revealed a similar binding constant for both fullerenes C₆₀ and C₇₀ (5×10^4 to 7×10^4 M⁻¹ in toluene), corroborated by DFT calculations that illustrate adaptive changes in nanohoop geometry upon fullerene complexation. ESI-MS is employed to generate ionized [1 : 1] host–guest complexes of diketo[9]CPP and DBP[9]CPP with C₆₀ and C₇₀ as guests. The relative stabilities of these complexes are evaluated in energy-resolved collision experiments.

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Introduction

[*n*]Cycloparaphenylenes (CPPs) are radially oriented π -conjugated nanohoops composed of *para*-linked benzene, possessing strikingly distorted phenylene moieties. Their unique physical and chemical properties make them promising materials for various fields, such as materials science, electronics, and biological applications.^{1–4} Various aromatic subunits other than benzene were introduced to tune the structural or optoelectronic properties of conjugated nanohoops beyond the size-dependent effect known for [*n*]CPPs.^{5–8} To date, a variety of nanohoops and their derivatives have been synthesized. These structures are a family of shape-rigid conjugated macrocycles featuring shape-persistent cavities surrounded by radially oriented π -systems. With these unique properties, conjugated nanohoops can be excellent hosts for fullerene guest molecules with high binding affinity.^{9–11} However, only a few examples have been reported to date on shape-adaptive nanohoops. Cong and coworkers reported a CPP-based figure-of-eight-shaped host enabling a similar guest-adaptive encapsulation of

fullerenes.¹² Stępień and coworkers published a squid-shaped molecule with a flexible cavity that binds neutral and cationic guests.¹³ Introducing flexible host structures can enhance guest selectivity and promote the dynamic binding of guest molecules, resulting in interesting host–guest properties and providing new insights into research areas such as guest recognition and separation.^{14–17}

Previously, we described the synthesis of diketo[6]CPP and diketo[7]CPP and the corresponding DBP[6]CPP and DBP[7]CPP (with R = mesityl, Fig. 1A).¹⁸ Diketo[*n*]CPPs are chiral based on the central chirality of the dihydroindeno[2,1-*a*]indene-5,10-dione unit (blue in Fig. 1A, structure in Fig. 1B). When the diketo units are transformed into dibenzo[*a,e*]pentalenes (DBPs), DBP[*n*]CPPs result, which are still chiral based on the bend in the DBP panel – as long as its rotation through the nanohoop is sterically hindered, which is the case for $n = 6, 7$, and also for $n = 8, 9$ reported herein. The enantiomers of diketo[*n*]CPPs can therefore be defined using (*S,S*)- or (*R,R*)-descriptors (Fig. 1B), while for the DBP[*n*]CPPs the M- and P-notations are suited.¹⁸ While both diketo[6]- and -[7]CPPs were highly fluorescent, leading to circularly polarized luminescence, their cavities were too small to host fullerene guest molecules. Building upon this work, we herein report on larger nanohoops diketo[8]- and diketo[9]CPPs as well as the derived DBP[8]- and DBP[9]CPPs, which we investigated for potential utilization in the fields of circularly polarized luminescence (CPL) materials and

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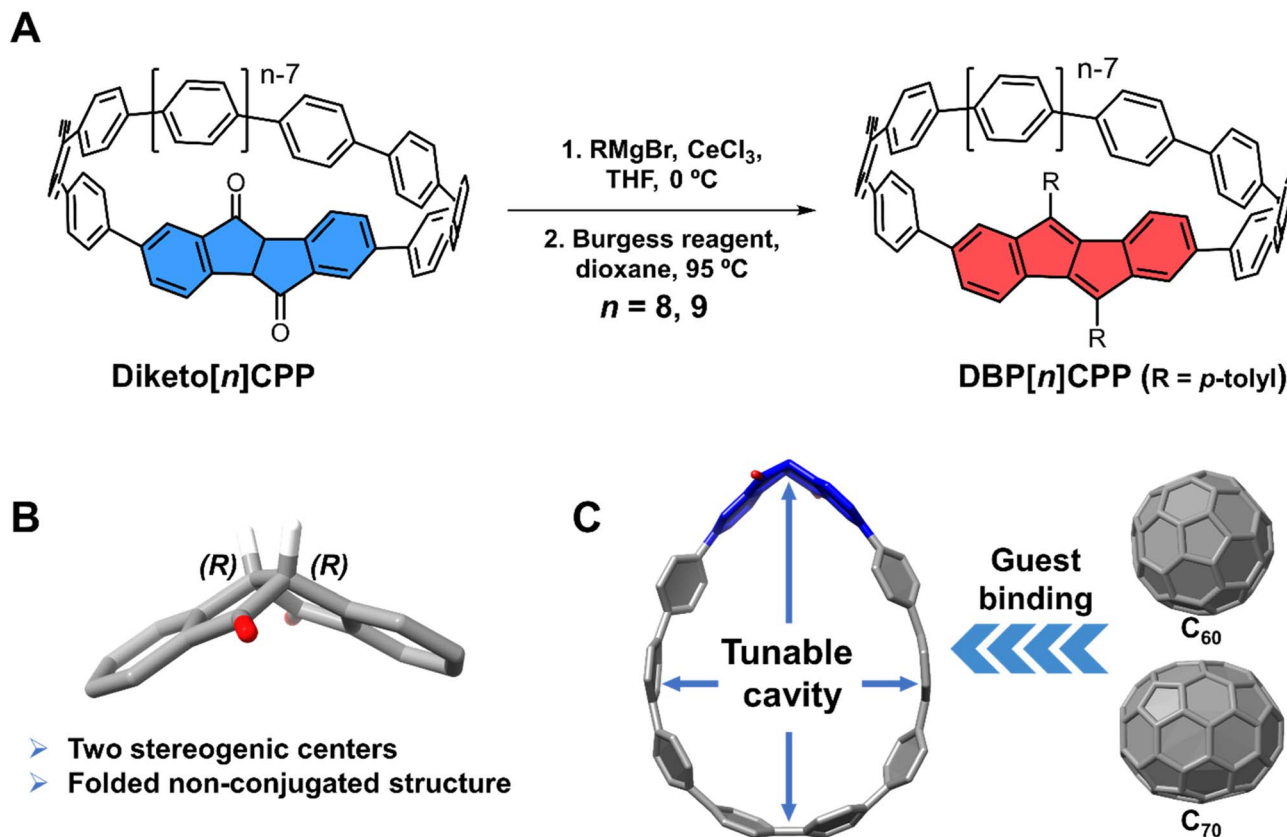


Fig. 1 (A) Diketone- and DBP-based chiral nano hoops synthesized herein; (B) unique structure of (*R,R*)-dihydroindeno[2,1-*a*]indene-5,10-dione as the central and chiral diketone unit; (C) the shape-adaptive cavity of diketo[8]- and diketo[9]CPPs enables binding of fullerenes. All structures were optimized with B3LYP-D3(BJ)/6-31G(d).

supramolecular chemistry. The cavities of diketo[8]- and diketo[9]CPP are large enough as well as shape-adaptive, so that they can both host C₆₀ or C₇₀ fullerene (Fig. 1C). Notably, the results of solution studies reveal that both diketo[*n*]CPPs have a similar binding affinity for C₇₀ and C₆₀, demonstrating that the molecules exhibit adaptive binding to these fullerenes. In addition, diketo[8]- and [9]CPPs show high fluorescence quantum yields of 0.87 and 0.92, respectively, which are enhanced compared to the smaller diketo[6]- and [7]CPPs (0.57 and 0.66, respectively).

Results and discussion

Synthesis

The synthesis of the diketo[*n*]CPPs and DBP[*n*]CPPs begins with the preparation of the key intermediates (Scheme 1). The bis-brominated kinked diketone **1** is a key building block in our synthetic route.¹⁹ In order to access larger diketo[*n*]CPPs in comparison to our previous studies,^{18,20} we extended this building block by attaching one (**2**, **3**) or two (**4**) further aryl units. The Suzuki–Miyaura coupling of dibromo–diketone **1** with 1.1 equivalents of 2-chlorophenylboronic acid afforded monoarylated diketone **2** in a yield of 75%. Then, the following Miyaura–borylation furnished compound **3** in a good yield of 77%. Separately, dibromo–diketone **1** was transformed into the π -extended bis(boronic ester) **4** by Pd-catalyzed borylation in

a yield of 46%. Subsequently, the macrocyclization of these intermediates (**3** and **4**) with C-shaped unit **5** (ref. 20 and 21) was carried out using a Pd-catalyzed Suzuki–Miyaura coupling reaction. This was followed by a reductive aromatization and furnished the larger diketo[*n*]CPPs in good yields of 22% (*n* = 8) and 30% (*n* = 9) over both steps.

In the next step, the diketo[*n*]CPPs were transformed into DBP[*n*]CPPs, for which we chose *p*-tolyl substituents. CeCl₃-mediated Grignard addition of *p*-tolyl magnesium bromide was followed by two-fold water elimination with the Burgess reagent as a mild dehydration method,^{19,22} which provided DBP[8]CPP and DBP[9]CPP in high yields of 88% and 92%, respectively, over these two steps.

DFT calculations at the B3LYP/6-31G(d) level of theory using the StrainViz²³ tool provided information on the strain energies of these nano hoops (see SI Section 9.1 for details). The strain energies amount to 42 and 41 kcal mol⁻¹ for diketo[8]CPP and diketo[9]CPP, and 59 and 53 kcal mol⁻¹ for DBP[8]CPP and DBP[9]CPP, respectively. These results show that the diketo[*n*]CPPs have lower strain energies than their corresponding DBP[*n*]CPPs, which is likely due to the naturally kinked structure of the diketone-units in the formed nano hoops, relieving part of the strain (see SI Table S54 and Fig. S63). Compared to the smaller diketo[6]CPP (strain energy 52 kcal mol⁻¹),¹⁸ diketo[8]CPP and diketo[9]CPP are significantly less strained. Diketo[*n*]CPPs and



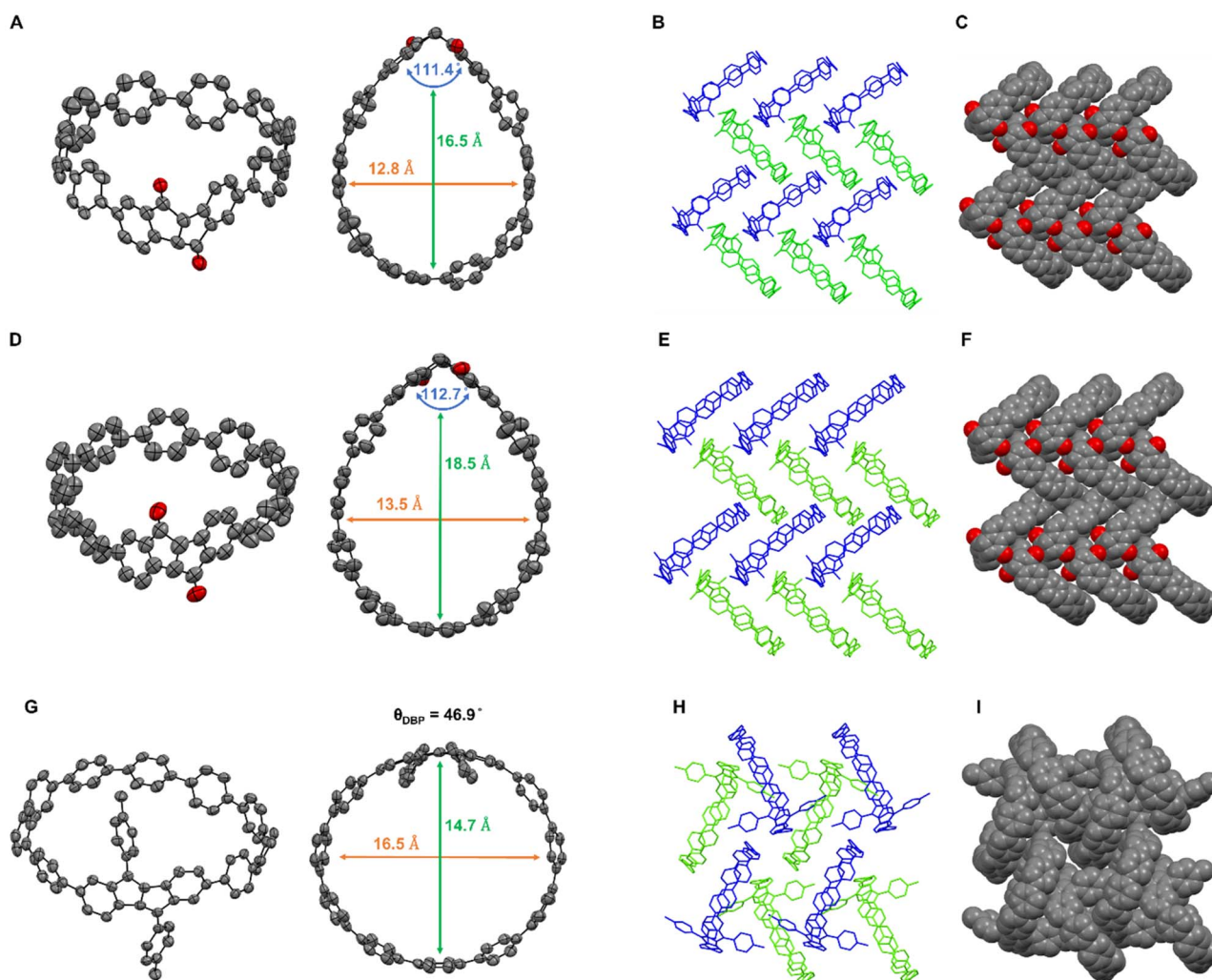


Fig. 2 Single-crystal X-ray structural analysis of diketo[8]CPP (A–C), diketo[9]CPP (D–F) and DBP[9]CPP (G–I). (A, D and G) Side view and top view of the crystal structures of diketo[9]CPP and DBP[9]CPP; (B, C, E, F, H and I) packing structure. (*S,S*) or (*R,R*) for the diketo[*n*]CPPs and *M*- or *P*- for the DBP[9]CPP enantiomers are highlighted in blue and green color, respectively. Solvent molecules and hydrogen atoms are omitted for clarity. Thermal ellipsoids are shown at the 50% probability level.

fluorescence spectroscopy (Fig. 3 and Table 1). The absorption spectra of diketo[*n*]CPPs and DBP[*n*]CPPs each show a strong band corresponding to the absorption of the oligoparaphenylene (PP) linkers as their maximum between 329 and 347 nm (Fig. 3A). Based on the results of TD-DFT (B3-LYP-D3(BJ)/def2-TZVP) calculations, these peaks correspond to multiple orbital transitions (see SI Tables S7–S10). For diketo[*n*]CPPs the main absorption bands appear at wavelengths close to diketo [6]- and [7]CPPs due to their similar orbital energy levels (Table 1 and SI Fig. S66). In addition, a weak shoulder absorption band at 360–420 nm can also be observed, which can be assigned to the HOMO → LUMO transition (SI Tables S7 and S8). The signature bands of the DBP units^{26–28} in the DBP[*n*]CPPs appear as peaks around 440–480 nm and a red-shifted long-wavelength-absorption onset compared to the diketo[*n*]CPPs. These peaks correspond to the HOMO → LUMO (first peak) and HOMO–1 → LUMO (second peak) transitions (see SI, Tables S9 and S10).

Compared to DBP[6]- and [7]CPPs, the DBP bands are more expressed in the larger DBP[8]- and [9]CPPs. The DBP[*n*]CPPs (*n* = 6–9) show higher molar attenuation coefficients (ϵ) than the diketo[*n*]CPPs. For all nano hoops, ϵ increases with the number of benzene rings (*n*), exhibiting a similar trend to [*n*]CPPs.² The molar attenuation coefficients ϵ exhibit a positive correlation with increasing hoop size and conjugation (see values in Table 1).

All four nano hoops are fluorescent with emission maxima from 437–460 nm (Fig. 3B, see SI Fig. S35 for DBP[8]- and [-9] CPPs, for exact values see Table 1). These emissions stem from the oligo(paraphenylene) moieties, which explains the similar wavelengths and shapes of the emission bands. Compared to the smaller congeners with *n* = 6 and 7,¹⁸ the emission maxima are blue-shifted both in the diketo- and DBP[8]- and [9]CPPs. This is in line with observations on [*n*]CPPs, where the emission blue-shifts with increasing size, a remarkable size-dependent



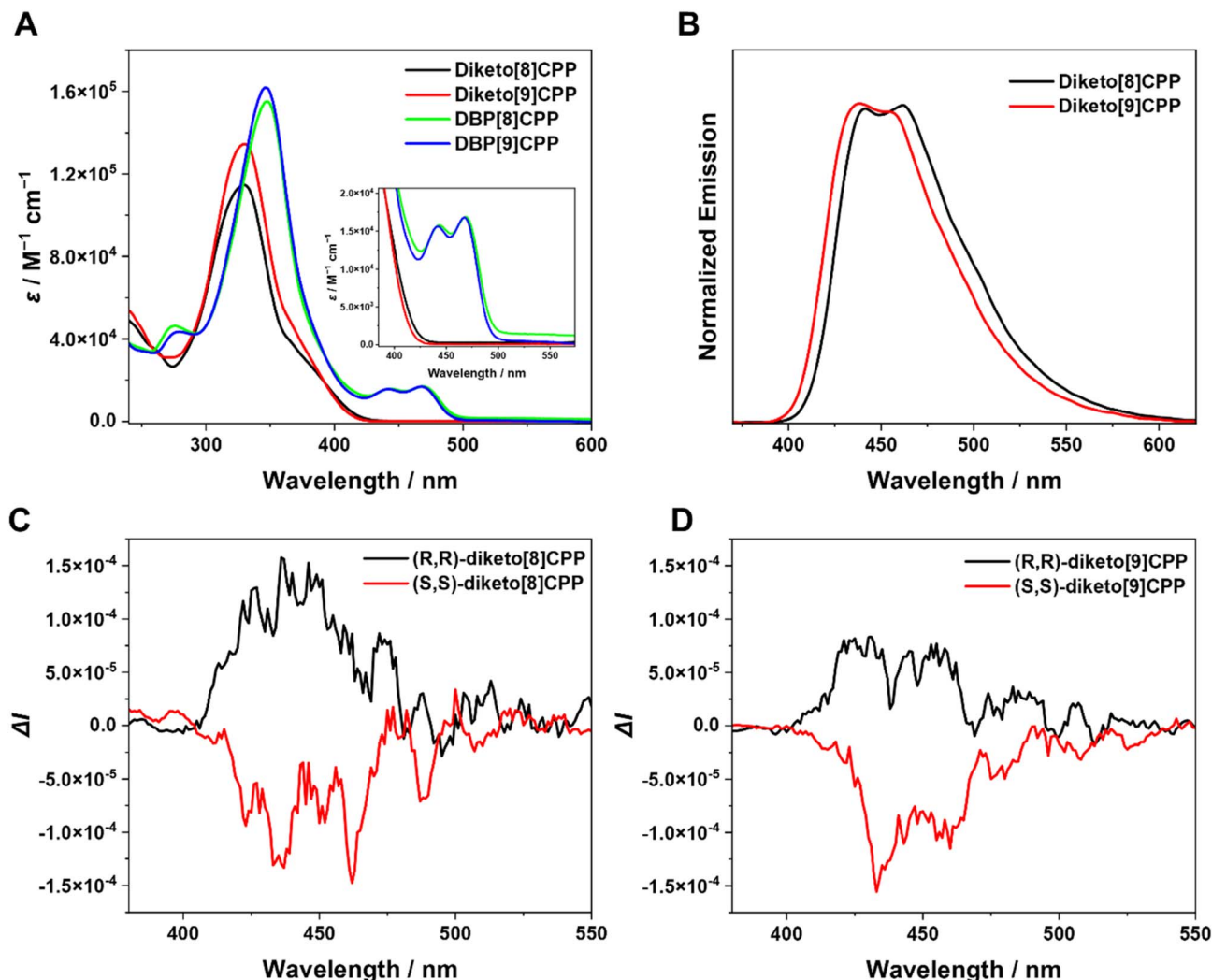


Fig. 3 (A) Absorption (10^{-5} M to 10^{-6} M) and (B) photoluminescence spectra (10^{-6} M to 10^{-7} M) of diketo[n]CPPs and DBP[n]CPPs; CPL spectra of (C) diketo[8]CPP and (D) diketo[9]CPP (5×10^{-5} M). All measurements were carried out in dichloromethane.

effect, which stands in contrast to linear conjugated PPs.¹ The multiple emission peaks likely arise from vibronic progressions of the lowest singlet excited state rather than from multiple emissive electronic states. Such vibronically resolved fluorescence is typical for [n]CPPs, where the 0–1/0–0 vibronic peak ratio has been explicitly analyzed²⁹ and strong electron–vibration coupling has been reported.³⁰ Remarkably, compared to

the smaller diketo[n]CPPs ($n = 6, 7$), diketo[8]CPP and diketo[9]CPP exhibited higher fluorescence quantum yields (Φ_F) of 0.87 and 0.92, respectively (for $n = 6$, $\Phi_F = 0.57$, and for $n = 7$, $\Phi_F = 0.66$).¹⁸ An increase in Φ_F with hoop size is also observed for [n]CPPs² and m [n]CPPs.²⁴ In contrast, DBP[8]CPP and DBP[9]CPP showed low Φ_F values (<0.01), which we ascribe to an energy or electron transfer occurring in the excited state from the PP units

Table 1 Optical properties^a and frontier molecular orbital energy levels of diketo[n]CPPs and DBP[n]CPPs

Compound	$\lambda_{\text{abs-max}}/\text{nm}$	$\epsilon/\text{M}^{-1} \text{cm}^{-1}$	$\lambda_{\text{em-max}}^c/\text{nm}$	Φ_F	$\Delta E_{\text{opt}}^e/\text{eV}$	$E_{\text{LUMO}}^d/\text{eV}$	$E_{\text{HOMO}}^d/\text{eV}$	$ g_{\text{lum}} $
Diketo[6]CPP ¹⁸	325	6.4×10^4	481	0.57	—	−1.98	−5.94	4.0×10^{-4}
Diketo[7]CPP ¹⁸	329	8.5×10^4	472	0.66	—	−2.01	−5.96	—
Diketo[8]CPP	329	1.1×10^5	460	0.87	2.97	−1.96	−5.95	6.8×10^{-4}
Diketo[9]CPP	330	1.3×10^5	437	0.92	3.01	−2.02	−5.95	6.9×10^{-4}
DBP[8]CPP ^b	347	1.5×10^5	437	<0.01	2.51	−2.46	−5.54	—
DBP[9]CPP ^b	346	1.6×10^5	452	<0.01	2.53	−2.45	−5.58	—

^a In dichloromethane. ^b Substituent: *p*-tolyl. ^c Excitation wavelength = 330 nm. ^d HOMO and LUMO levels calculated at the PW6B95-D3(BJ)/def2-QZVPP level of theory. ^e Determined from the onset of the longest wavelength absorption.



to the non-emissive DBP part of the compounds, which quenches their fluorescence. Pentalene derivatives have been shown to undergo rapid decay from the S_1 state *via* singlet fission.³¹ Structurally similar monobenzopentalenes also show no S_1 emission, which has been attributed to a formally forbidden $S_1 \rightarrow S_0$ transition and fast intersystem crossing or singlet fission.³² Emission has only been observed from the S_2 state, with extremely low quantum yields (<0.1%). Consequently, the DBP units are non-emissive and dissipate the energy radiation-free.²⁷ Hence, the optical properties of the diketo[n]CPPs are dominated by the PP linkers with excellent fluorescence quantum yields. In DBP[n]CPPs, the optical properties reflect the presence of both the PP linkers in characteristic absorption and emission bands as well as that of the DBP units in a red-shifted absorption onset and an almost complete quenching of the fluorescence intensity.

Next, we studied the chiroptical properties of the nano-hoops. The (*S,S*)- and (*R,R*)-enantiomers of diketo[8]- and [9]CPPs were separated using HPLC on a chiral stationary phase. The CD and CPL spectra for all compounds are shown in Fig. 3C or SI Section 6. Each pair of enantiomers displays mirror-imaged CD and CPL spectra with opposite Cotton effects. Time-dependent density functional theory (TD-DFT) calculations indicate that the theoretical CD spectra for diketo[n]CPPs are compatible with the experimental results (Fig. S38), from which we assigned the enantiomers. In the excited state, all enantiomers exhibit distinct CPL spectra, with dissymmetry factors $|g_{lum}|$ at λ_{em} (max) reaching 6.8×10^{-4} and 6.9×10^{-4} for diketo[8]CPP and diketo[9]CPP, respectively—values slightly higher than those reported for diketo[6]CPP.¹⁸ Although these dissymmetry factor values fall within the typical range for chiral organic emitters, the relatively high molar extinction coefficient and quantum yield of the two keto-nanohoops result in valuable CPL brightness values (B_{CPL})³³ of 32.5 and $41.2 \text{ M}^{-1} \text{ cm}^{-1}$ for diketo[8]CPP and diketo[9]CPP, respectively.

We then attempted to separate the enantiomers of DBP[8]CPP, however, unsuccessfully. Re-injection of an initially pure enantiomer of DBP[8]CPP stored at room temperature for several minutes already revealed the presence of the opposite enantiomer, indicating rapid racemization under ambient conditions. Furthermore, enantiomeric separation of DBP[9]CPP showed a slight plateau between the two peaks in the HPLC elugram. This chromatographic behavior was observed even at a column-oven temperature as low as $20 \text{ }^\circ\text{C}$, indicating a low racemization barrier (Fig. S39). Thus, the racemization behavior of DBP[9]CPP was investigated by means of dynamic variable-temperature (VT) HPLC on a chiral stationary phase.³⁴ The elution profiles of DBP[9]CPP in the temperature range from 283 to 303 K in steps of 5 K were recorded and analyzed with the DCXplorer software packages (see SI Section 6.2).³⁵ With increasing column temperature, the retention time separation between the two enantiomers became smaller, indicating a faster exchange rate. The data were then fitted using the Eyring equation to give the thermodynamic parameters $\Delta H_e^\ddagger = 18.6 \text{ kcal mol}^{-1}$ and $\Delta S_e^\ddagger = -31.6 \text{ J mol}^{-1}$. Therefore, the inter-conversion energy barrier between the enantiomers was estimated to be $20.9 \text{ kcal mol}^{-1}$ at 298 K , which is close to the DFT-

calculated barrier ($19.9 \text{ kcal mol}^{-1}$). For DBP[8]CPP, the elution profile was then analyzed *via* the unified equation for dynamic chromatography,³⁵ which enabled the determination of a racemization barrier of $\Delta G_e^\ddagger = 23.5 \text{ kcal mol}^{-1}$ at 308 K , consistent with the DFT barrier of $22.5 \text{ kcal mol}^{-1}$, explaining its markedly slower racemization. We previously reported calculated racemization barriers for methyl-substituted DBP[6]CPP and DBP[7]CPP of $36.5 \text{ kcal mol}^{-1}$ and $26.9 \text{ kcal mol}^{-1}$, respectively. These smaller nanohoops exhibit higher barriers and slower racemization.¹⁸

Supramolecular host properties of diketo[n]CPPs

With diameters between 13.5 \AA and 16.5 \AA , both diketo[8]- and diketo[9]CPPs as well as DBP[8]- and DBP[9]CPPs have cavities suitable to host C_{60} and C_{70} fullerenes. The host-guest complexes formed between the nanohoops and both C_{60} and C_{70} were characterized by high-resolution mass spectrometry (HRMS), NMR and fluorescence spectroscopy. ^1H NMR spectra of diketo[n]CPPs and DBP[n]CPPs in $\text{TCE-}d_2$ showed clear shifts of the proton resonances upon addition of excess C_{60} or C_{70} (see SI Section 7.2), except for the addition of C_{60} to DBP[9]CPP. This might be due to the large size of this nanohoop resulting in a reduced interaction with the C_{60} fullerene. All signals in the complexes show chemical shift changes or splitting, which can be rationalized by their different spatial locations relative to the shielding and deshielding regions created by the fullerene guests. HRMS analyses indicated ionized species with m/z values matching diketo[n]CPPs $\supset C_{60}$ or DBP[n]CPPs $\supset C_{70}$ for all nanohoops, demonstrating the possibility of host-guest-complex formation, also for DBP[9]CPP (see SI Section 7.1).

We quantified the binding of diketo[n]CPPs to fullerenes by conducting fluorescence titration experiments in toluene solution. The fluorescence of the diketo[n]CPPs is significantly quenched upon addition of C_{60} or C_{70} (Fig. 4A and B and SI Section 7.3), suggesting that a charge transfer may occur between the host and guest, similar to [n]CPPs.³⁶ This was supported by the calculated frontier molecular orbitals of diketo[n]CPPs $\supset C_{60}$ and diketo[n]CPPs $\supset C_{70}$, where the HOMO is localized on the diketo[n]CPP and the LUMO on the fullerene (SI, Fig. S69). By fitting the titration curves using the BindFit program,^{37–40} we obtained the binding constants (K) for the host-guest complexes in toluene at room temperature. The average association constants for diketo[8]CPP $\supset C_{60}$, diketo[8]CPP $\supset C_{70}$, diketo[9]CPP $\supset C_{60}$ and diketo[9]CPP $\supset C_{70}$ are 6.3×10^4 , 7.2×10^4 , 5.2×10^4 , and $5.3 \times 10^4 \text{ M}^{-1}$, respectively, in toluene (Table 2). These are lower than the binding constant for the C_{60} complex of [10]CPP in toluene of $2.8 \times 10^6 \text{ M}^{-1}$,³⁶ determined by fluorescence quenching titration, but similar to that for the C_{60} complex of [4]cyclo-dibenzopentalene of $5.3 \times 10^4 \text{ M}^{-1}$ in toluene (from UV/vis-spectroscopic titration).⁴¹ The binding constants of the diketo[n]CPPs lie much closer to that of [10]CPP with C_{70} fullerene of $8.4 \times 10^4 \text{ M}^{-1}$ in toluene.⁸ This is a very interesting observation: while [10]CPP binds C_{60} more than 30 times stronger than C_{70} in solution, for each diketo[n]CPP we obtained similar binding constants for both fullerenes ($K(C_{60})/K(C_{70}) \approx 1$). We reason that this observation is related to



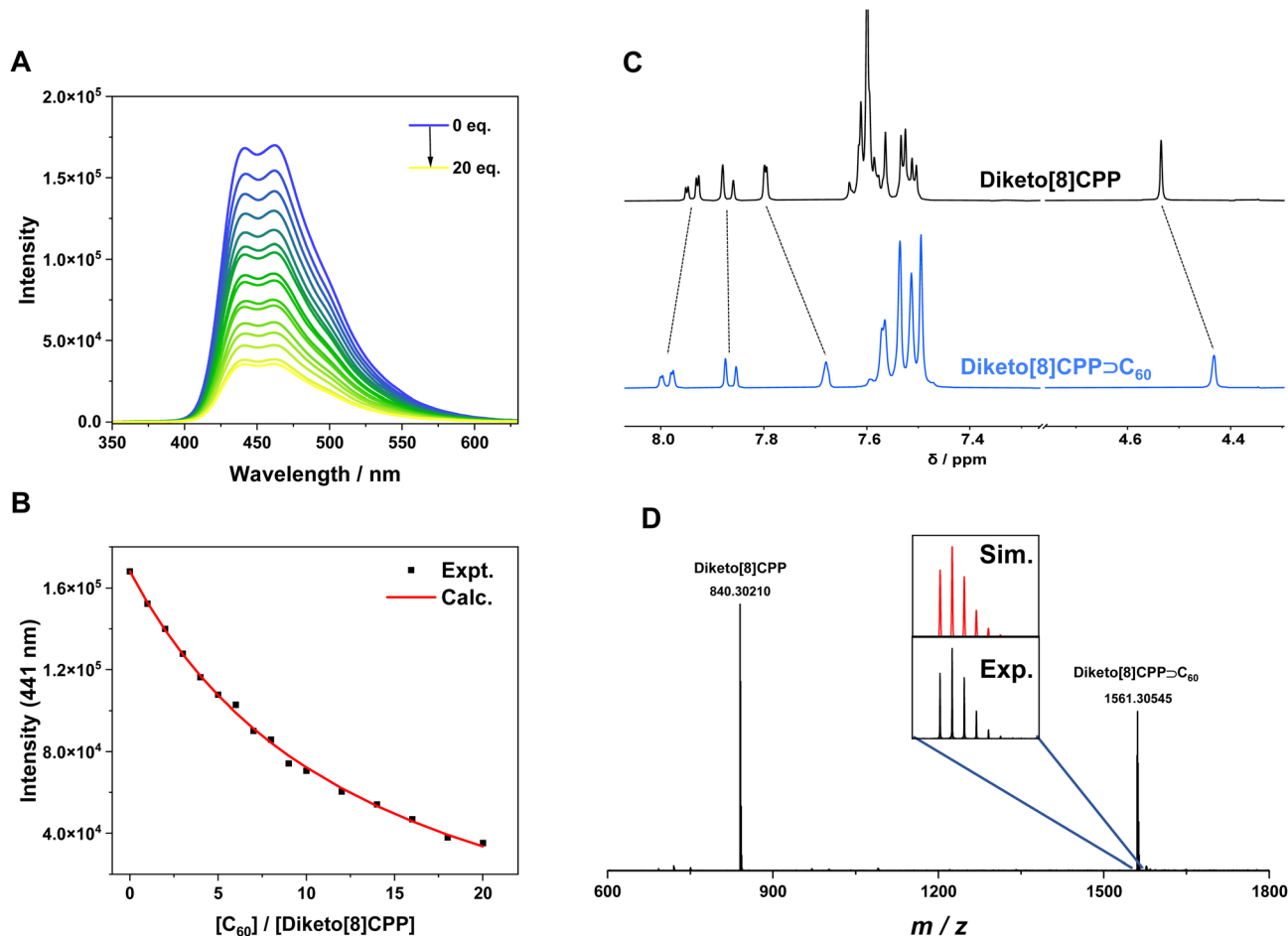


Fig. 4 Supramolecular host properties of diketo[*n*]CPPs. (A and B) Fluorescence titration of diketo[8]CPP (1.2×10^{-6} M) with C_{60} in toluene (fit based on a 1 : 1 model); (C) ^1H NMR spectra (1,1,2,2-tetrachloroethane- d_2 , 400 MHz) of diketo[8]CPP and its mixture with 3 eq. C_{60} ; (D) MALDI mass spectrum of diketo[8]CPP $\supset C_{60}$ (inset shows the experimental and calculated isotopic pattern).

Table 2 Experimental binding constants^a K of 1:1 host–guest complexes and free binding energies ΔG (298 K) from fluorescence titrations of diketo[*n*]CPPs with C_{60} and C_{70} in toluene

Experiment		K^a/M^{-1}	$\Delta G^b/\text{kcal mol}^{-1}$
Diketo[8]CPP	C_{60}	$6.27(\pm 0.02) \times 10^4$	−6.55
	C_{70}	$7.15(\pm 0.04) \times 10^4$	−6.62
Diketo[9]CPP	C_{60}	$5.20(\pm 0.02) \times 10^4$	−6.43
	C_{70}	$5.33(\pm 0.01) \times 10^4$	−6.45

^a Average of three titrations. ^b Calculated using $\Delta G = -RT \ln K$.

the shape adaptivity of the diketo[*n*]CPPs, maximizing their interacting surface with the fullerene guest in each case.

To further assess the non-covalent interactions and shape adaptivity of the diketo[*n*]CPPs towards fullerene binding, we optimized their molecular geometries as well as those of the fullerene complexes at the B3LYP-D3(BJ)/6-31G* level of theory (Fig. 5 and Table 3). C_{60} is approximately spherical with a mean diameter of 7.1 Å and fits well into both diketo[8]- and -[9]CPPs. It preferably interacts with the π -surfaces of the PP parts, especially visible in diketo[9]CPP $\supset C_{60}$, where it is centered at

the “bottom” part of the nanohoop as shown in Fig. 5. The nanohoops adapt their shape to best host C_{60} : while the “kink” angle θ of the diketo unit in diketo[8]CPP increases from 110.7° to 114.9° in the fullerene complex, this angle gets smaller in diketo[9]CPP (111.6° to 108.3°) to maximize interactions of the π -surfaces. Diketo[8]CPP overall assumes a more circular shape (*a* and *b* values approaching each other) to host C_{60} .

Because of the anisotropic shape of fullerene C_{70} , reminiscent of an American football with a short axis of 7.12 Å and a long axis of 7.96 Å, the supramolecular complexes formed between nanohoops and C_{70} are diverse.^{8,42} Based on the calculated association energies, C_{70} adopts a standing orientation in diketo[8]CPP, but a lying orientation in diketo[9]CPP (Fig. 5 and Table S11). Similar to the C_{60} complex, diketo[8]CPP has to widen its “kink” angle θ to 114.3° to best host C_{70} , while for diketo[9]CPP accommodating C_{70} in standing orientation hardly requires any geometrical change of the nanohoop ($\theta = 110.2^\circ$). As for C_{60} , diketo[8]CPP assumes a more circular shape in the C_{70} complex.

For each of the fullerene complexes, an independent gradient model based on Hirshfeld partition (IGMH) analysis^{43,44} showed that the *para*-phenylene moieties around the



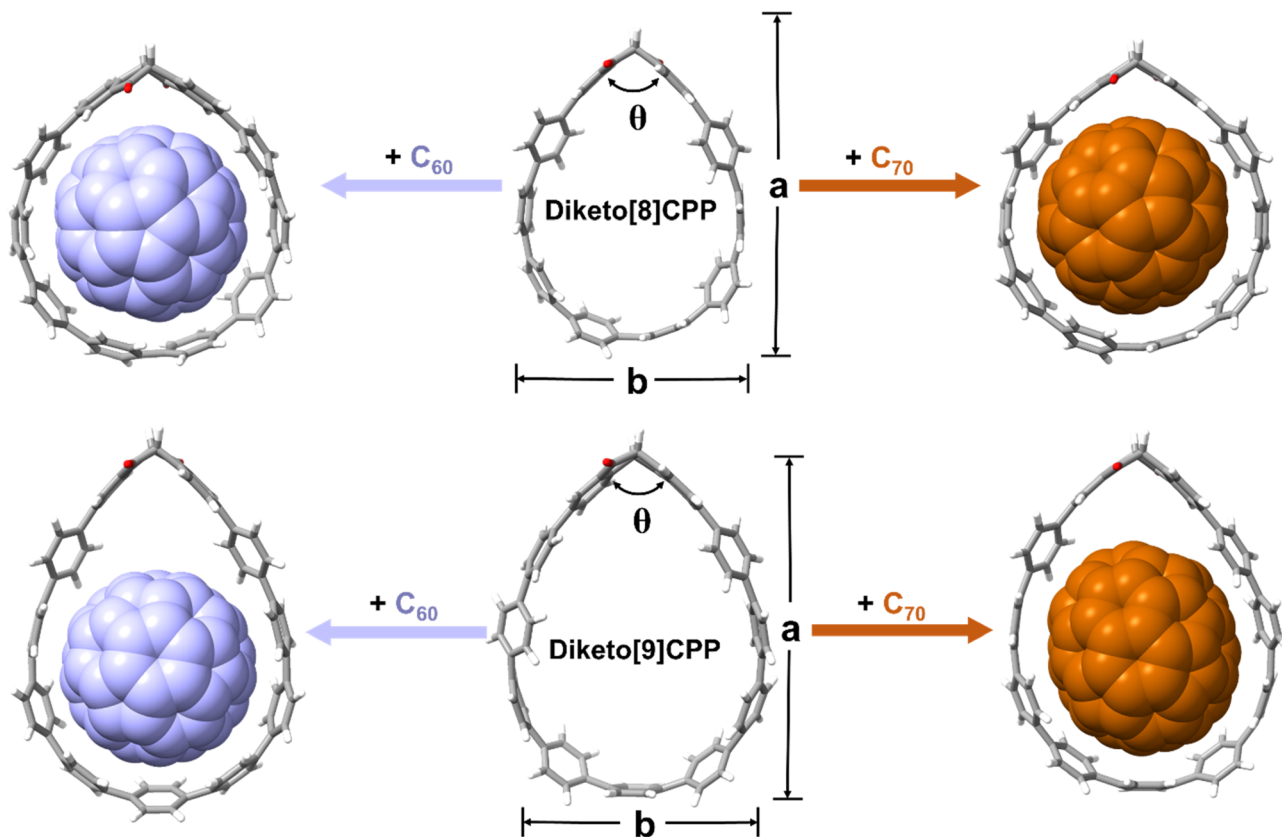


Fig. 5 DFT-optimized geometries (B3LYP-D3(BJ)/6-31G(d)) and association free energies (GFN2-xTB) of diketo[n]CPPs and their fullerene complexes.

Table 3 Geometrical parameters of diketo[n]CPPs and their fullerene complexes, including long axis (a), short axis (b), and bending angle (θ)

	$a/\text{\AA}$	$b/\text{\AA}$	$\theta/^\circ$
Diketo[8]CPP	16.66	12.29	110.7
Diketo[8]CPP \supset C ₆₀	15.31	13.83	114.9
Diketo[8]CPP \supset C ₇₀	15.31	13.84	114.3
Diketo[9]CPP	18.22	13.43	111.6
Diketo[9]CPP \supset C ₆₀	18.30	13.48	108.3
Diketo[9]CPP \supset C ₇₀	18.0	13.96	110.2

fullerene guest in each cavity contribute to the non-covalent interactions (SI, Fig. S70).

The calculated association-free energies⁴⁵ for diketo[8]CPP \supset C₆₀, diketo[8]CPP \supset C₇₀, diketo[9]CPP \supset C₆₀ and diketo[9]CPP \supset C₇₀ at the GFN2-xTB level of theory in toluene amount to -5.5 , -7.6 , -4.7 and -5.6 kcal mol⁻¹, respectively (see SI Table S13). This compares well to the experimental results in toluene (Table 2).

Single crystals of the complex diketo[9]CPP \supset C₆₀ suitable for X-ray diffraction were obtained by vapor diffusion of MeCN into a 1,2-dichlorobenzene solution. The solid-state structure clearly shows that a complex between diketo[9]CPP and C₆₀ with 1 : 1 stoichiometry is present. The fullerene inside the cavity was restrained and placed in two positions (see SI Section 8.2). The top half of the structure is disordered over two positions

consisting of both enantiomers of the diketo[9]CPP structure with a ratio of 57/43. In the solid state, incorporation of C₆₀ transforms the herringbone packing of diketo[9]CPP (Fig. 2E and F) into a columnar structure. Notably, the overall size of the diketo[9]CPP cavity shows only a slight variation upon complexation (Fig. 6).

The above-discussed solution-based complexation experiments were complemented by gas-phase experiments on the formation and relative stability of noncovalent complexes of diketo[9]- and DBP[9]CPPs with C₆₀ and C₇₀, respectively. For this, electrospray ionization-(tandem) mass spectrometry (ESI-(MS/MS)) was applied. ESI is a soft ionization method allowing the gentle transfer of loosely bound complexes from solution into the gas phase. In the present case, the formation of ions occurs by electrochemical oxidation with the ESI source operating as an electrochemical cell. Consequently, radical cations are formed through the oxidation of the molecules that constitute the noncovalent complex. The resulting ionized complexes are subjected to energy-resolved collision-induced dissociation (CID). Recording the decay as a function of the applied collision energy allows establishing the relative bond strength holding together the complex ion. The E_{50} value refers to the collision energy, at which half of the complex has decomposed, and is taken as a measure of stability. All oxidized nanohoop–fullerene complexes decomposed into the nanohoop radical cation by loss of a neutral fullerene. Evidently, within



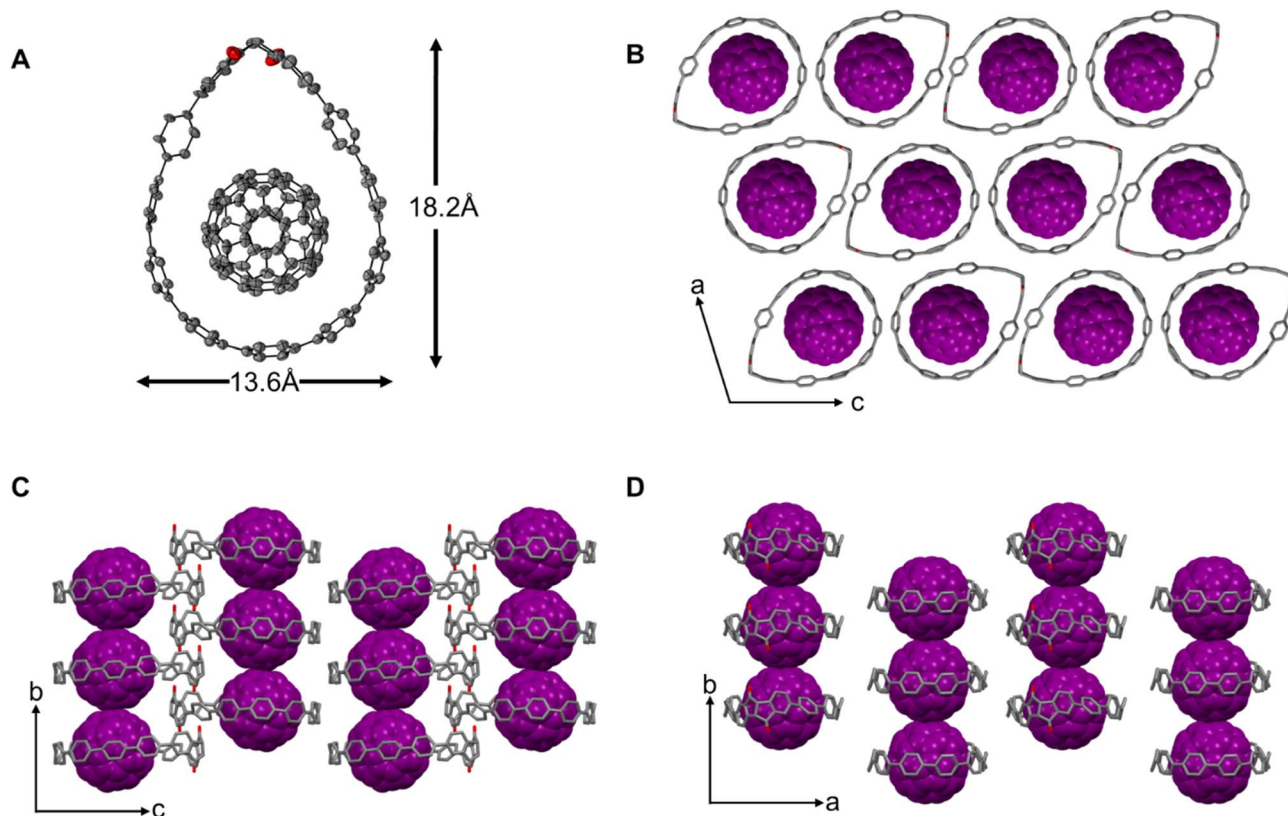


Fig. 6 Single-crystal X-ray structural analysis of diketo[9]CPP \supset C₆₀. (A) Top view of the solid-state structure of diketo[9]CPP \supset C₆₀. (B–D) Packing structure. Solvent molecules and hydrogen atoms are omitted for clarity. Thermal ellipsoids are shown at the 50% probability level.

the complex, the nano hoop was oxidized in favour of the fullerene.

Table 4 compares the E_{50} values as a measure of the relative stabilities of the cationic host–guest complexes combining [n]CPP ($n = 10, 11$) as well as diketo[9]- and DBP[9]CPPs as the hosts with C₆₀ and C₇₀ as guests. [10]CPP⁺⁺⊃C₆₀ shows the strongest noncovalent bonding amongst the C₆₀ complexes.^{36,46} [11]CPP⁺⁺⊃C₆₀ is slightly less strongly connected. The C₆₀ complexes of diketo[9]- and DBP[9]CPP⁺⁺ show only a slightly lower stability as [11]CPP⁺⁺⊃C₆₀, with the DBP[9]CPP complexes being marginally more stable than the diketo[9]CPP complexes. The stability of all host–guest complexes increases when C₇₀ is incorporated. Even though [10]CPP has a considerably larger binding constant with C₆₀ than with C₇₀ in equilibrium measurements in solution,^{8,36} the dissociation of the isolated complex ions in the gas phase reveals a stronger bonding to C₇₀. The latter has been observed before and

attributed to enhanced interactions due to the larger π system of C₇₀.⁴⁶ [11]CPP⁺⁺⊃C₇₀ is the most stable C₇₀ complex followed by [10]CPP⁺⁺⊃C₇₀. Also diketo[9]CPP, whose shape-adaptive binding led – in solution – to similar binding constants as C₆₀ and C₇₀, forms more stable isolated complex ions with C₇₀ than with C₆₀. Diketo[9]CPP⁺⁺⊃C₇₀ is slightly more stable than DBP[9]CPP⁺⁺⊃C₇₀. Evidently, the host–guest complexes employing the shape-adaptive diketo[9]CPP gain similar stabilities as those using DBP[9]CPP and [11]CPP as the hosts.

Conclusions and outlook

In summary, we presented diketo[n]CPPs and DBP[n]CPPs with $n = 8$ and 9 and studied their structures, photophysical properties, and interactions with fullerenes. All four nano hoops are chiral, with diketo[n]CPPs based on the central chirality of the diketo unit, and DBP[n]CPPs based on the hindered rotation of the DBP panels. Remarkably, diketo[n]CPPs showed high fluorescence quantum yields of 87% ($n = 8$) and 92% ($n = 9$), while the fluorescence in DBP[n]CPPs was quenched, likely due to energy transfer to the non-emissive DBP units. Diketo[8]CPP and diketo[9]CPPs deliver valuable CPL brightness values (B_{CPL}) of 32.5 and 41.2 M⁻¹ cm⁻¹, respectively, while the DBP[n]CPPs racemize easily owing to their lower enantiomerization barriers. The diketo unit bestows diketo[n]CPPs with a shape-adaptive structure, where through the change of its kink angle the

Table 4 E_{50} values of the investigated noncovalent radical cation complexes (host \supset guest¹⁺⁺) in eV

Guest	Host			
	[10]CPP	[11]CPP	Diketo[9]CPP	DBP[9]CPP
C ₆₀	0.146	0.117	0.099	0.103
C ₇₀	0.158	0.172	0.141	0.115



geometrical parameters of the nanohoops can adapt. This unique property enabled diketo[8]- and diketo[9]CPPs to form stable complexes with both C₆₀ and C₇₀, showing similar binding strengths in the range of 5×10^4 to 7×10^4 M⁻¹ in toluene to both fullerene guests due to their shape adaptivity. This is supported by DFT calculations on the structures of the inclusion complexes, clearly demonstrating the changes in the nanohoop shape. Collision experiments in the gas phase reveal that noncovalent complex ions with C₆₀ and C₇₀ possess similar individual stabilities using either diketo[9]CPP or DBP[9]CPP as the host with bond strengths similar to the use of [11]CPP as the host. Having uncovered the basic structural and optoelectronic properties of diketo[n]CPPs and DBP[n]CPPs, we are now in a position to fully utilize these unique nanohoops for various applications.

Author contributions

Conceptualization: B. E. and X. X.; data curation: all authors; formal analysis: X. X., M. H., L. Y., P. S., and L. E.; funding acquisition: B. E.; investigation: X. X., M. H., L. Y., P. S., and L. E.; methodology: X. X., M. H., L. Y., P. S., and L. E.; project administration: B. E.; supervision: B. E., G. P., and T. D.; validation: all authors; visualization: X. X.; writing – original draft: X. X. and B. E.; writing – review & editing: all authors.

Conflicts of interest

There are no conflicts to declare.

Data availability

The data that support the findings of this study are openly available in Zenodo at <https://zenodo.org/doi/10.5281/zenodo.15847822>, reference number 15847822.

CCDC 2364437, 2365575, 2427456 and 2486379 contain the supplementary crystallographic data for this paper.^{47a-d}

Supplementary information: materials and methods, synthetic procedures and characterization data, NMR and mass spectra, further spectra from UV/Vis, fluorescence, CD and CPL measurements, details on host-guest studies, SC-XRD data, and details on DFT calculations. See DOI: <https://doi.org/10.1039/d5sc05305f>.

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