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1,2-Hydrogen atom transfer of aminyl radicals under photoredox catalysis for the synthesis of α -amino phosphine oxides

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α -Amino phosphorus compounds are significant structural motifs in natural products, pharmaceuticals and organocatalysis. The construction of such motifs by hydrogen atom transfer (HAT), however, is rare. Herein, we describe a photocatalytic net 1,2-HAT of nitrogen-centered radicals for the formation of C(sp³)-P bonds. This net 1,2-HAT proceeds *via* deprotonation of N-centered radicals without the requirement of external strong bases, instead being facilitated by the benzoate anion generated during the reaction. This process enables the synthesis of α -amino phosphine oxides bearing various functional groups under mild conditions. The gram-scale synthesis demonstrates the scalability of the net 1,2-HAT rearrangement/radical-radical coupling reaction, which has been used in the preparation of an antitumor agent. Mechanistic investigations, including radical trapping experiments, Stern-Volmer fluorescence quenching experiments, cyclic voltammogram studies, cross-over experiments, KIE studies and DFT calculations support a net 1,2-HAT pathway.

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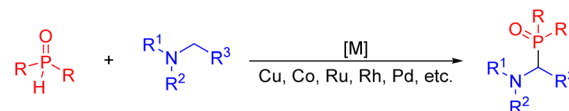
Introduction

Organophosphorus compounds are common structural motifs in natural products and play vital roles in pharmaceuticals and organocatalysis.¹⁻⁶ As a subclass of these, α -amino phosphonates and their derivatives are of interest in the pharmaceutical and agricultural industries, because they possess unique biological activities allowing uses as herbicides,⁷ cancer cell inhibitors,⁸ and virus inhibitors.⁹⁻¹² Moreover, α -amino phosphorus compounds can be efficient catalysts¹³ in organic synthesis and they are also used as P,N ligands^{14,15} in organometallic chemistry and catalysis. As such, more efficient and eco-friendly methods to prepare α -amino phosphorus compounds remain in demand.

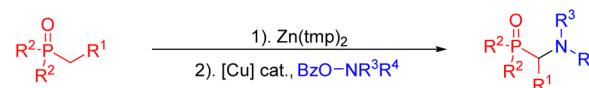
Given the utility of α -amino phosphorus compounds, various approaches have been developed for their preparation. The Kabachnik-Fields reaction^{16,17} (hydrophosphonylation of imines), which involves the reaction of an amine, an aldehyde

or a ketone, and a phosphinic acid, represents one of the most common methods toward α -amino phosphorus compounds. Alternatively, the α -phosphonylation of amines to prepare α -amino phosphorus compounds has been demonstrated with various transition-metal catalysts, including Cu, Co, Rh, Ru and

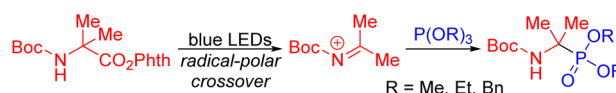
a. Transition-metal-catalyzed α -phosphonylation of amines



b. Copper-catalyzed amination of α -phosphonate zincates

Wang, et al. *ACIE*, 2014

c. Photocatalytic decarboxylative phosphorylation

Aggarwal, et al. *ACIE*, 2022

Scheme 1 Representative approaches to the synthesis of α -amino phosphorus derivatives. (a) Transition-metal-catalyzed α -phosphonylation of amines. (b) Copper-catalyzed amination of α -phosphonate zincates. (c) Photocatalytic decarboxylative phosphorylation.

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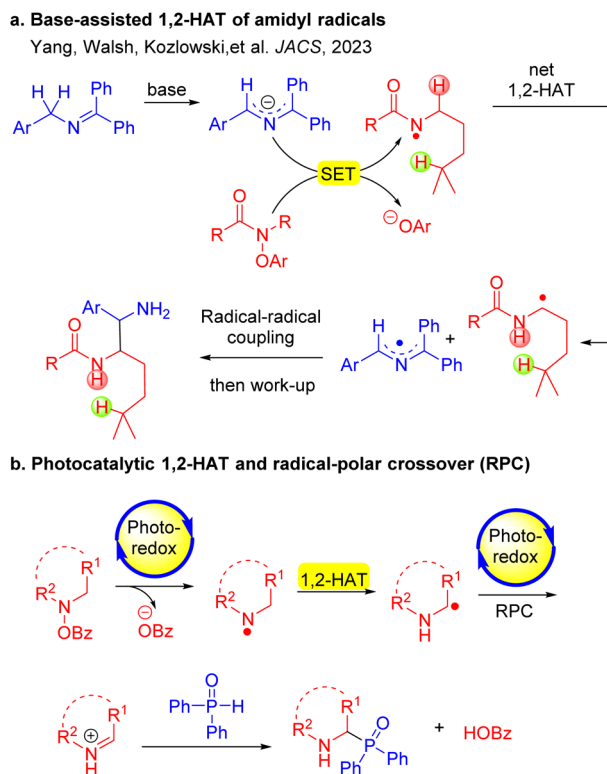
Pd (Scheme 1a). In some cases, these reactions are conducted with high temperatures and pressures.^{18–26} In 2014, Wang and co-workers developed an approach to α -aminophosphine oxides using copper-catalyzed electrophilic α -amination of α -phosphonate zincates to form C–N bonds (Scheme 1b).²⁷ Recently, progress has been made in the formation of C–P bonds through catalytic reactions employing light and electrochemistry.^{28–32} For example, Aggarwal's team utilized photoredox catalysis^{33–37} to facilitate the decarboxylation of α -amino acid derivatives. The resulting alkyl radicals were oxidized to *N*-acyliminium ions, ultimately being captured using triphenyl phosphite (Scheme 1c).³⁸

Hydrogen atom transfer (HAT) processes are strategically important for achieving regioselective functionalization of C–H bonds. Heteroatom (O,N)-centered radicals, which generally undergo intramolecular 1,5-HAT reactions (Hofmann–Löffler–Freitag reaction),³⁹ produce translocated C-centered radicals. 1,2-HAT is less frequently observed compared to the intramolecular 1,5-HAT process. This is because the 1,5-HAT process is favored by a chair-type six-membered ring transition state, resulting in a lower energy⁴⁰ requirement for hydrogen atom migration. Recently, there have been a limited number of reports highlighting the net 1,2-HAT of O-centered radicals (O \cdot) in the literature.^{41–43} Notably, 1,2-HAT reactions of nitrogen-centered radicals (N \cdot) exhibit great potential in synthesis, but are rarely observed. Recently, our team developed a method to

coax amidyl radicals to undergo net 1,2-HAT processes to deliver α -amino carbon-centered radicals (C \cdot) (Scheme 2a). The newly formed C-centered radicals coupled with 2-azaallyl radicals^{37,44–46} to generate 1,2-diamine derivatives. Mechanistic experiments and density functional theory (DFT) calculations revealed that this 1,2-HAT is a base-assisted, stepwise process (Scheme 2a).⁴⁷ Subsequently, the Chen group reported base-promoted 1,2-HAT processes to prepare fluorinated amines.^{48,49}

Considering the significance of α -amino phosphorus compounds, we hypothesized that our newly developed 1,2-HAT of nitrogen-centered radicals (N \cdot) could lead to a new approach to form C–P bonds. Previous reports on the reaction between diphenylphosphine oxide and *O*-acylhydroxylamines resulted in P–N bond formation to afford phosphinic amides [R₂P(O)NR₂].^{50,51} However, to the best of our knowledge, there are no reports on reactions of these coupling partners to access α -amino phosphorus derivatives.

Herein, we report a photocatalytic net 1,2-HAT of nitrogen-centered radicals for the formation of C(sp³)-P bonds (Scheme 2b). Diverging from former net 1,2-HAT protocols, the present approach uniquely employs the benzoate anion—functioning dually as a leaving group and a base promoter—to facilitate the reaction. Specifically, hydroxyl amines were found to undergo a photocatalytic single electron transfer (SET) process to generate N-centered radicals.⁵² The N-centered radical undergoes a benzoate-promoted isomerization through a net 1,2-HAT process to generate a C-centered α -amino radical, which undergoes SET with the photoredox catalyst (4CzIPN) to form the iminium ion in a radical-polar crossover (RPC). A phosphine oxide participates in an intermolecular coupling with the resultant iminium to furnish α -amino phosphorus products in good yields. This method enables the synthesis of α -amino phosphine oxides bearing various functional groups (37 examples, up to 88% yield). Mechanistic investigations, including radical trapping experiments, Stern–Volmer fluorescence quenching experiments, cyclic voltammogram studies, cross-over experiments, KIE studies and DFT calculations, support a net 1,2-HAT pathway. Notably, this net 1,2-HAT reaction takes place employing a simple organic photoredox catalyst without the addition of base or transition metals.



Scheme 2 1,2-HAT of nitrogen-centered radicals. (a) Base-assisted net 1,2-HAT of amidyl radicals. (b) Photoredox initiated net 1,2-HAT of nitrogen-centered radicals followed by radical-polar crossover (this work).

Results and discussion

We selected 4-benzoyloxy-morpholine (**1a**) as the initial substrate because the morpholine group is common in bioactive molecules. Reaction of **1a** was performed with diphenylphosphine oxide **2a** in the presence of photoredox catalysts (2 mol%) in PhCF₃ at room temperature under irradiation with 450 nm blue LEDs for 24 h. Initially, a series of commercial photoredox catalysts including iridium, ruthenium and organic photoredox catalysts were evaluated (Table 1, entries 1–7). Among them, the organic photoredox catalyst 4CzIPN generated the α -amino phosphine oxide **3aa** in 79% assay yield (AY, as determined by ¹H NMR integration against an internal standard, entry 7), while other iridium and ruthenium photoredox catalysts resulted in product **3aa** in lower AY



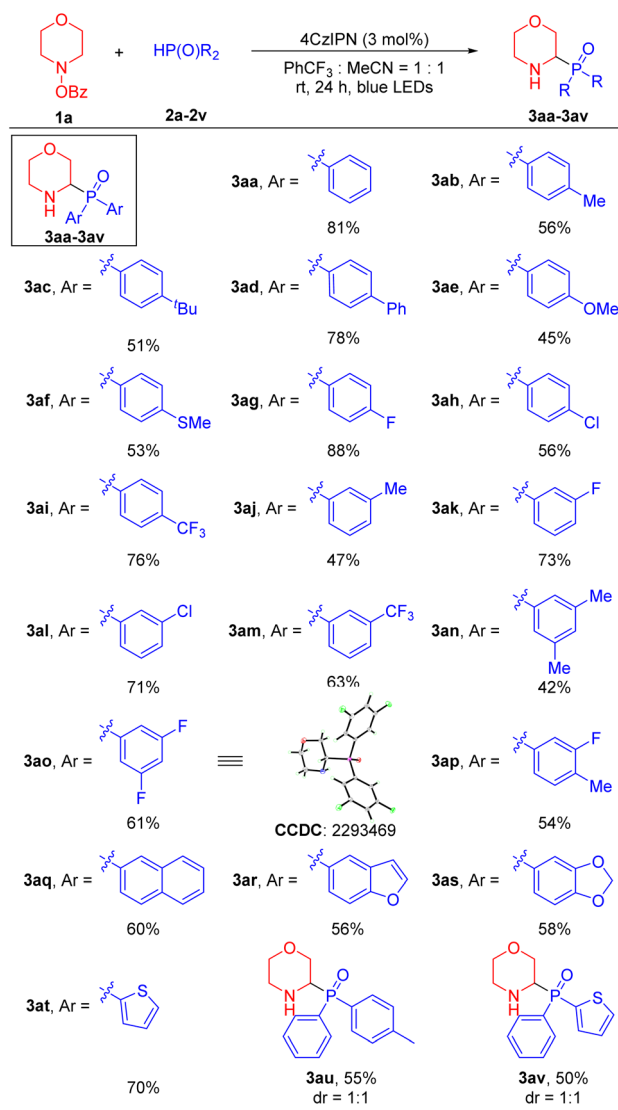
Table 1 Optimization of the reaction conditions^{a,b}

Entry	Photocatalyst	Solvent	Yield (%)
1	Ir(ppy) ₂ (bpy)PF ₆	PhCF ₃	Trace
2	Ir(ppy) ₃	PhCF ₃	0
3	Ir(<i>p</i> -tBu-ppy) ₃	PhCF ₃	72
4	Ir[dF(CF ₃)ppy] ₂ (dtbpy)PF ₆	PhCF ₃	48
5	Ir(dtbbpy)(ppy) ₂ PF ₆	PhCF ₃	69
6	Ru(bpy) ₃ (PF ₆) ₂	PhCF ₃	47
7	4CzIPN	PhCF ₃	79
8	4CzIPN	THF	0
9	4CzIPN	DCM	35
10	4CzIPN	DCE	64
11	4CzIPN	MeCN	50
12	4CzIPN	DMF	0
13	4CzIPN	DMSO	0
14	4CzIPN	Toluene	Trace
15	4CzIPN	MeCN : PhCF ₃ = 1 : 1	83
16	4CzIPN (1 mol%)	MeCN : PhCF ₃ = 1 : 1	51
17	4CzIPN (3 mol%)	MeCN : PhCF₃ = 1 : 1	85 (81)^c
18	4CzIPN (4 mol%)	MeCN : PhCF ₃ = 1 : 1	82
19	—	MeCN : PhCF ₃ = 1 : 1	0
20 ^d	4CzIPN	MeCN : PhCF ₃ = 1 : 1	0

^a Reaction conditions: **1a** (0.2 mmol, 2.0 equiv.), **2a** (0.1 mmol, 1.0 equiv.), photoredox catalyst (2 mol%), rt, 24 h, under blue LEDs (450 nm). ^b Assay yields determined by ¹H NMR spectroscopy of the unpurified reaction mixtures using CH₂Br₂ as an internal standard. ^c Isolated yield. ^d Without blue LEDs.

(up to 72%) or no reaction. Using 4CzIPN as photoredox catalyst, we examined a series of solvents including THF, DCE, MeCN, DMF, DMSO and toluene (entries 8–14). However, these solvents led to lower AY (up to 64%) or no reaction. To improve the solubility of 4CzIPN, we tested mixed solvent systems (see the SI for details). Among them, the reaction performed best in 1 : 1 MeCN : PhCF₃ (83% AY, entry 15). Reducing the photocatalyst loading from 2 to 1 mol% afforded only 51% AY (entry 16). In contrast, raising the photoredox catalyst to 3 mol% provided the desired product **3aa** in 85% AY and 81% isolated yield (entry 17). Conducting the reaction using 4 mol% photoredox catalyst led to a slight decrease to 82% AY (entry 18). Finally, control experiments (entries 19 and 20), which confirmed that both photoredox catalyst 4CzIPN and blue LED irradiation are required to form product **3aa**. Based on these results, the optimal conditions for the net 1,2-HAT are those in entry 17 of Table 1.

With effective α -amino phosphorylation conditions identified (Table 1, entry 17), we evaluated the scope of the phosphine oxide substrates. As shown in Table 2, a wide range of diarylphosphine oxides with different aryl groups delivered the α -amino phosphine oxides in moderate to good yields. Diarylphosphine oxides with electron-donating groups at the 4-position, such as 4-Me (**2b**), 4-^tBu (**2c**), 4-Ph (**2d**), 4-OMe (**2e**) and 4-SMe (**2f**), reacted with 4-benzoyloxy-morpholine **1a** to generate

Table 2 Scope of phosphine oxides **2**^{a,b}

^a Reactions were conducted on a 0.4 mmol scale using 2.0 equiv. **1a**, 1.0 equiv. **2** and photoredox catalyst (3 mol%) at 0.1 M. ^b Yields of isolated products after chromatographic purification.

α -amino phosphorylation products **3ab**, **3ac**, **3ad**, **3ae** and **3af** in 45–78% yields, respectively. Diarylphosphine oxides with electronegative and electron-withdrawing groups at the 4-position, such as 4-F (**2g**), 4-Cl (**2h**) and 4-CF₃ (**2i**) furnished products **3ag**, **3ah** and **3ai** in 88%, 76% and 56% yields, respectively. Diarylphosphine oxides with groups at the 3-position, such as 3-Me (**2j**), 3-F (**2k**), 3-Cl (**2l**) and 3-CF₃ (**2m**) led to products **3aj**, **3ak**, **3al** and **3am** in 47–63% yields. Diarylphosphine oxides containing disubstituted aryl groups, such as 3,5-di-Me (**2n**), 3,5-di-F (**2o**) and 3-F-4-Me (**2p**) gave products **3an**, **3ao** and **3ap** in 42–61% yields. Furthermore, the structure of product **3ao** was confirmed by X-ray crystallography (CCDC 2293469; see the SI for details). The π -extended 2-naphthyl substrate (**2q**) afforded product **3aq** in 60% yield. It is noteworthy that this method was



also well-suited for medicinally relevant heterocyclic derivatives. Diarylphosphine oxides bearing benzofuranyl (**2r**), piperonyl (**2s**) and 2-thiophenyl (**2t**) substituents furnished the corresponding products **3ar**, **3as** and **3at** in 56–70% yields. Finally, when diarylphosphine oxides containing two different aryl groups (**2u** and **2v**) were employed, the products **3au** and **3av** were formed in 55% and 50% yields, respectively, with no diastereoselectivity (*dr* = 1 : 1).

Next, we evaluated the ability of the net 1,2-HAT radical reaction to accommodate various hydroxylamine derivatives, which were easily synthesized from the corresponding amines and benzoyl peroxide^{53–55} (see the SI for details). Generally, a wide range of cyclic and acyclic amines were compatible with our method, generating the α -amino phosphine oxides in moderate to good yields (Table 3). Cyclic amines including pyrrolidine (**1b**), piperidine (**1c**), thiomorpholine (**1d**), hexamethyleneimine (**1e**), 4,4-dimethylpiperidine (**1f**), 4-piperidone ethylene ketal (**1g**) and *cis*-octahydroisindole (**1h**), reacted with diphenylphosphine oxide **2a** to afford the desired products **3ba**,

3ca, **3da**, **3ea**, **3fa**, **3ga** and **3ha** in 47–81% yields. Furthermore, acyclic dialkylamino groups, such as dimethylamine (**1i**), diethylamine (**1j**), isopropylmethylamine (**1k**), dibutylamine (**1l**), *N*-methyl-cyclohexylamine (**1m**) and *N*-methyl-4-amino-tetrahydropyran (**1n**), were also suitable partners, delivering the corresponding products **3ia**, **3ja**, **3ka**, **3la**, **3ma** and **3na** in 43–60% yields. It is noteworthy that in the cases of **3ka** and **3am** the aminyl radical intermediate undergoes deprotonation preferentially at the least sterically hindered position, giving the higher energy primary radical over the more stable tertiary radical. Finally, in order to showcase the synthetic utility of our approach and further probe selectivity in the net 1,2-HAT, we carried out the phosphorylation of druglike molecules. For instance, hydroxylamine derivatives from antidepressant nortriptyline^{56,57} (**1o**) and fluoxetine^{58,59} (**1p**) underwent net 1,2-HAT reactions to afford products containing the α -phosphorylation fragment in **3oa** and **3pa** in 43% and 45% yields, respectively. Here again, selective deprotonation of the α -aminyl radical at the methyl group over the secondary C–H positions was observed.

To test the scalability of this net 1,2-HAT radical coupling process, we carried out a gram-scale synthesis (Scheme 3a). Reaction of 4-benzoyloxy-morpholine **1a** (10 mmol) was performed with diphenylphosphine oxide **2a** (5 mmol) under the standard conditions for 24 hours to provide 1.09 g of the target product **3aa** with a yield of 76%. This result is comparable to the 81% yield achieved on a 0.1 mmol scale (Table 1). To establish the synthetic utility of our method, we employed a two-step transformation to synthesize the antitumor agent **4ba**.⁶⁰ First, α -amino phosphine oxide **3ba** was afforded in 47% yield under the standard conditions. Then, **3ba** was treated with benzenesulfonyl chloride under basic conditions to generate the active **4ba** in 72% yield. The synthesis of this bioactive product illustrates the practical application of our method to prepare pharmaceutically relevant substances (Scheme 3b).

To elucidate the reaction mechanism, we undertook a series of investigative experiments. First, a radical trapping

Table 3 Scope of hydroxylamine derivatives **1**^{a,b}

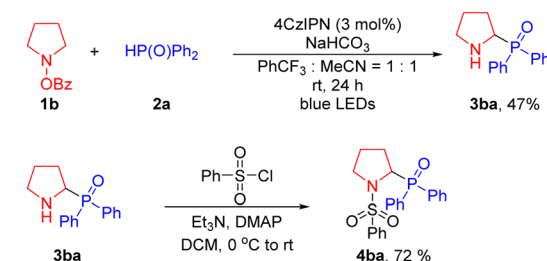
	47% ^[b]	
	72%	
	56%	
	58%	
	63%	
	81%	
	58%	
	52% ^[b]	
	53% ^[b]	
	60%	
	43% ^[b]	
	56%	
	59%	
	43% ^[b]	from Nortriptyline
	45% ^[b]	from Fluoxetine

^a Reactions were conducted on a 0.4 mmol scale using 2.0 equiv. **1**, 1.0 equiv. **2a** and photoredox catalyst (3 mol%) at 0.1 M. ^b Yield of isolated product after chromatographic purification. ^c 3.0 equiv. **1** was used and 3.0 equiv. NaHCO₃ was used.

a. Gram-scale synthesis of **3aa**.

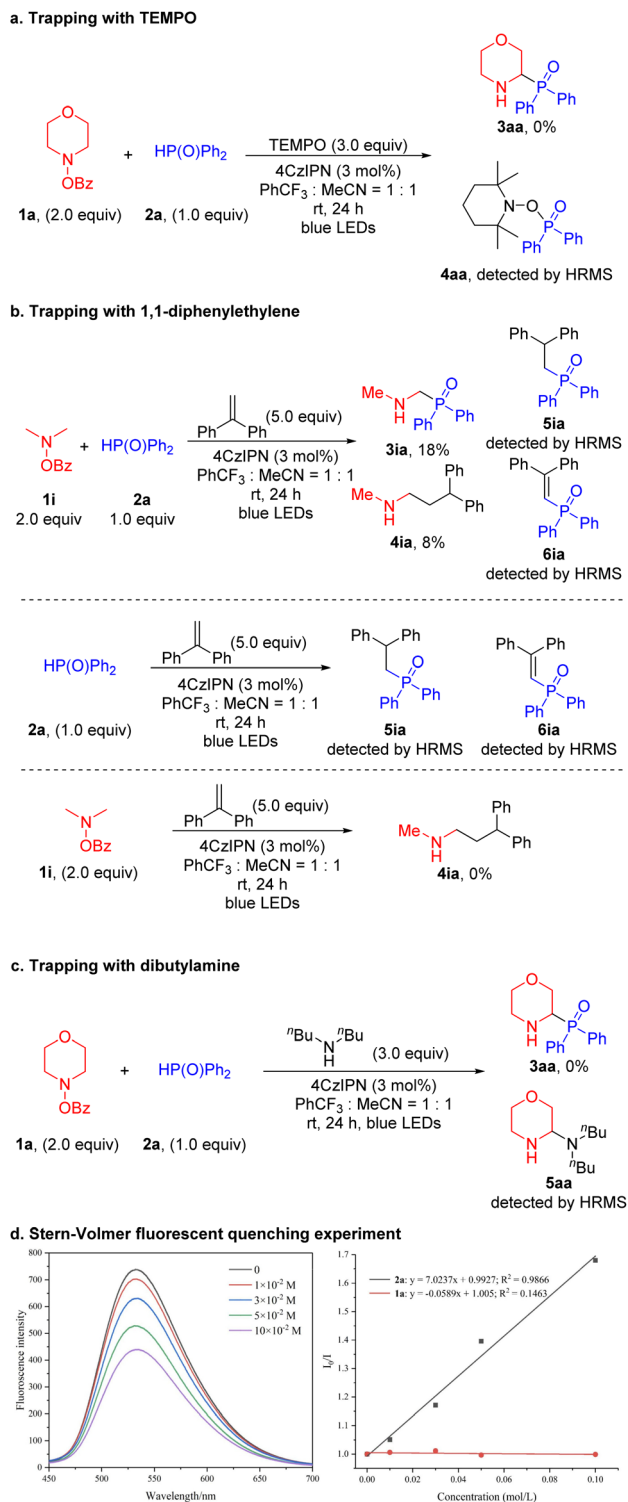


b. Synthesis of antitumor agent **4ba**.



Scheme 3 Gram-scale synthesis and product derivatization. (a) Gram-scale synthesis of **3aa**. (b) Preparation of an antitumor agent (**4ba**).





Scheme 4 Mechanistic studies. (a) Trapping with TEMPO. (b) Trapping with 1,1-diphenylethylene. (c) Trapping with dibutylamine. (d) Stern–Volmer fluorescent quenching experiments. Quenching of the 4CzIPN emission (4.0×10^{-4} M in MeCN:PhCF₃ = 1:1) in the presence of increasing amounts of 2a.

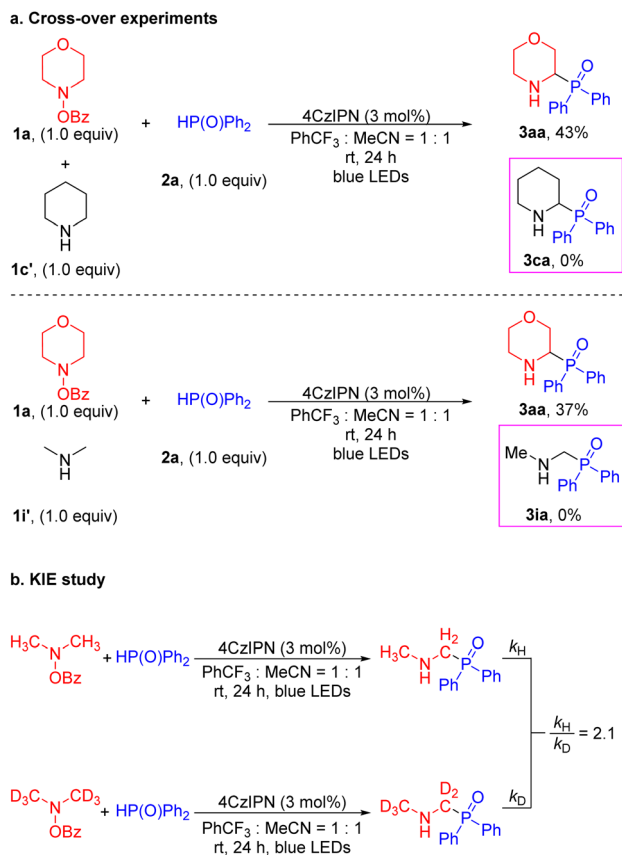
experiment using 2,2,6,6-tetramethyl piperidine-1-oxyl (TEMPO, 3.0 equiv.) was conducted with 1a and 2a as substrates (Scheme 4a). The reaction failed to generate the product 3aa and the

oxidized phosphine oxide trapped with TEMPO (4aa) was detected by HRMS. Product 4aa is not stable to silica gel and decomposed during isolation. Additionally, when 2a was used and no 1a was added, 4aa was again detected (see the SI for details). Based on the work of Doyle, we conducted trapping experiments using 1i and 2a as substrates, along with the radical trap 1,1-diphenylethylene (5.0 equiv.) (Scheme 4b).⁶¹ The yield of the expected product 3ia was only 18%. Trace products formed from addition of a phosphorus-based intermediate to the 1,1-diphenylethylene (5ia and 6ia) were detected by HRMS, further supporting generation of a P-centered radical. Given that only trapping of the P-centered radical is observed and no trapped α -amino radical is detected, we hypothesize that the photoredox catalyst initiates the process through oxidation of the phosphine oxide (as discussed below). We also isolated the product formed from the α -amino radical addition to diphenylethylene (4ia, 8% yield). The reaction between 2a and 1,1-diphenylethylene under the standard conditions, but in the absence of amine, also led to coupling products 5ia and 6ia (detected by HRMS). However, the reaction between 1i and 1,1-diphenylethylene in the absence of phosphine oxide did not generate the trapping product 4ia. This result supports the photoredox catalyst first oxidizing the phosphorus coupling partner and then reducing the amine derivative. Furthermore, when dibutylamine, a nucleophile with low reactivity toward radical reactions, was added, the corresponding aminal compound 5aa was generated and detected by HRMS (Scheme 4c). This result supports the presence of an iminium cation as an intermediate.

To further explore this mechanistic pathway, a series of Stern–Volmer quenching experiments were performed (Scheme 4d). These experiments also revealed that the excited state of 4CzIPN* was quenched by the phosphorus coupling partner rather than the amine derivative (see the SI for details). Additionally, the photoredox catalyst 4CzIPN with an $E_{1/2}(^*P/P^-)$ of +1.35 V (vs. SCE) can oxidize 2a ($E_{ox} = +1.0$ V vs. SCE).^{62,63} On the other hand, the photoredox catalyst 4CzIPN with an $E_{1/2}(P^+/P^*)$ of -1.04 V (vs. SCE)⁶² is unlikely to reduce 1a ($E_{re} = -2.04$ V vs. SCE, see the SI for more details). This result is also consistent with a photocatalytic cycle initiated with the oxidation of 2a.

It is known that the α -amino C–H bonds are generally weaker than other C–H bonds and the resulting radicals are stabilized relative to isolated C-centered radicals.^{47,64–68} With this in mind, we probed whether the HAT process was an intra- or intermolecular reaction by conducting cross-over experiments (Scheme 5a). In the event, treatment of a 1:1 mixture of 4-benzoyloxy-morpholine 1a and piperidine 1c' in the presence of diphenylphosphine oxide 2a under the standard conditions led to the morpholine-derived product 3aa (43% yield) without the generation of the cross-over piperidinyll-containing product 3ca (Scheme 5a). Similarly, reaction of 4-benzoyloxy-morpholine 1a and dimethylamine 1i' with 2a furnished the product 3aa (37% yield) without the cross-over product 3ia. The cross-over experiments verified that the generation of the α -amino radical was an intramolecular process. Moreover, additional evidence was obtained by examining the parallel kinetic isotope effect (KIE). A KIE value of 2.1 was measured, consistent with removal of the α -

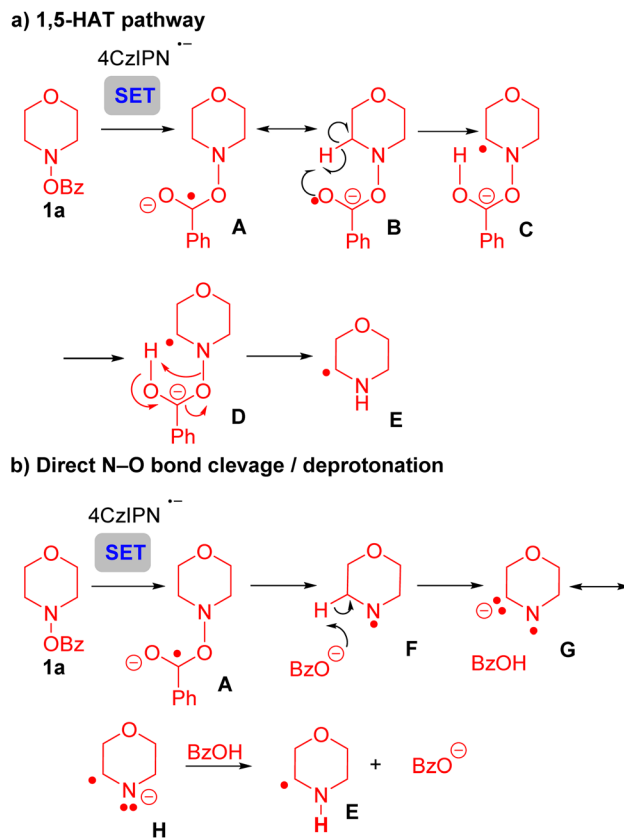




Scheme 5 Mechanistic experiments. (a) Cross-over experiments. (b) KIE study.

amino C–H/D contributing to the rate-limiting step (Scheme 5b).⁶⁹ This observation leads us to propose a low barrier for cleavage of the N–O bond.

To further probe the mechanism of this transformation, we turned to density functional theory (DFT) calculations [UM06/6-311+G(d,p)-CPCM(acetonitrile)//UB3LYP/6-31G(d),^{70–76} see SI for full computational details]. Given the KIE of 2.1 (Scheme 5) we initially considered two mechanisms that could account for these results. Initially, we explored the formation of product **3aa** via a 1,5-HAT process (Scheme 6a and Fig. S12). Here, it was envisioned that the reduced photocatalyst underwent SET with the *N*-benzyloxy substrate **1a** to give the reduced intermediate **A** and its resonance form **B**. Intermediate **B**, with radical character on the oxygen could be envisioned to undergo a 1,5-HAT to give **C**. Intermediate **C** might undergo N–O bond cleavage with proton transfer to generate the α -amino radical. While the transition state for the transfer of a hydrogen atom from the α -carbon directly to the benzoate oxygen could be located with the bonds frozen, the removal of the constraints led to the N–O bond being cleaved directly after SET (Fig. S13). We attribute this observation to the lability of the N–O bond as well as the unfavorable, rigid nature of the 1,5-HAT transition state. Similarly, transfer of the hydrogen atom to the nitrogen with the N–O bond intact (**C-TS-E**, Fig. S14) could not be located without the bonds constrained. Thus, the 1,5-HAT pathway was discounted.



Scheme 6 Overview of the mechanism. (a) 1,5-HAT pathway. (b) Direct N–O bond cleavage/deprotonation.

The second pathway, a net 1,2-HAT process, yields product **3aa**, as outlined in Scheme 6b. Following SET from the photocatalyst to substrate **1a**, radical anion **A** is again formed. N–O bond rupture of **A** gives the N-centered radical **F** and benzoate. Deprotonation of radical **F** by benzoate generates the resonance stabilized anion **G/H**. Reprotonation at nitrogen forms the α -amino radical **E**.

The energetics of a direct 1,2-HAT (Fig. S15) vs. base assisted net 1,2-HAT (Scheme 6a) were compared using DFT calculations. Both processes begin with species **1a'** undergoing facile N–O cleavage (Scheme 7, top, **A-TS-F**, 1.7 kcal mol⁻¹) to generate N-centered radical **F** and the benzoate anion (–44.8 kcal mol⁻¹). Dissociation of the benzoate anion followed by direct-1,2-HAT via the 3-centered transition state (**F1-TS-VII**, Fig. S15) has a high barrier of 42.8 kcal mol⁻¹ as expected.⁴⁸ Consequently, we explored the formation of the C-centered radical **VII** by an indirect, base-assisted-1,2-HAT, similar to our previous findings.⁴⁸ In this pathway (Scheme 7, bottom), the benzoate anion deprotonates the C–H bond alpha to the N-centered radical via **F-TS-G** (barrier of 23.7 kcal mol⁻¹) to form radical **G**. Subsequent transfer of the H atom to the nitrogen via **G-TS-E** is facile (barrier of 6.5 kcal mol⁻¹) and exergonic to form the benzoate anion and C-centered radical **E** (–47.0 kcal mol⁻¹). Dissociation of the benzoate anion leads to α -amino C-centered radical intermediate **VII**, which can subsequently undergo SET



because it is facilitated by a very weak base, specifically the benzoate counteranion released upon N–O cleavage. This observation is instructive in that it underscores the high acidity of the aminyl radical α -C–H bonds. Unsymmetrically substituted nitrogen-centered radicals show excellent selectivity in the HAT, which enables the synthesis of α -amino phosphine oxides bearing various functional groups (37 examples, up to 88% yield). A gram-scale synthesis demonstrates the scalability of the net 1,2-HAT coupling reaction. The synthesis of antitumor-active molecules highlights the promising potential applications of this reaction. Mechanistic investigations, including radical trapping experiments, Stern–Volmer fluorescence quenching experiments, cyclic voltammogram studies, cross-over experiments, KIE experiments, and DFT calculations support a net 1,2-HAT radical pathway followed by a radical polar crossover. It is also noteworthy that this net 1,2-HAT reaction involves a simple organic photoredox catalyst at room temperature, and without the addition of transition metals. The sustainability of the approach makes it an attractive choice for applications in the pharmaceutical industry.

Author contributions

X. Y. conceived of the project. H. Z., B. L. and P. J. W. designed the experiments. A. P., Y. M., X. T., S. D. and Y. J. performed the research. M. E. R. and M. C. K. performed DFT calculations. X. Y., P. J. W. and M. C. K. wrote the manuscript.

Conflicts of interest

There are no conflicts to declare.

Data availability

CCDC 2293469 contains the supplementary crystallographic data for this paper.⁷⁸

All experimental data, procedures for data analysis, and pertinent data sets are provided in the supplementary information (SI). Supplementary information is available. See DOI: <https://doi.org/10.1039/d5sc00268k>.

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