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Metal-decorated Graphdiyne Nanocarriers for Favipiravir Delivery: A DFT Investigation

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Abstract:

Favipiravir (FAV) is a broad-spectrum antiviral whose therapeutic efficacy is limited by poor bioavailability, motivating the development of nanoscale delivery platforms. Herein, density functional theory (DFT) calculations are employed to systematically investigate the adsorption behavior of FAV on pristine and transition-metal-decorated graphdiyne (GDY) nanosheets. Ni, Cu, and Ti modifiers are introduced to tune interfacial interactions at the atomic scale. Calculations at the B3LYP/6-31G(d) level, in both gas and implicit solvent environments, reveal that pristine GDY enables moderate, reversible physisorption (−0.26 eV; 2.45 Å), suitable for controlled release. In contrast, metal decoration significantly strengthens adsorption, with FAV(N/O)–GDY–Ni, –Cu, and –Ti exhibiting binding energies of −4.41, −3.96, and ~~−6.03~~ −5.79 eV, respectively, alongside reduced interaction distances. Electronic structure analyses (FMO, DOS, NBO, and RDG-NCI) confirm enhanced charge transfer and interaction localization upon metal incorporation, supported by thermodynamic favorability. These findings highlight metal-decorated GDY as a tunable nanoplatform for improving drug loading and stability, offering computational framework for guiding future experimental work.

Keywords: Favipiravir drug, Graphdiyne nanosheet, Drug delivery, Decoration, Nanocarriers, DFT study.

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1. *Introduction*

Favipiravir (FAV) was originally developed and approved for the treatment of influenza virus infections (1). Favipiravir is an antiviral drug that selectively inhibits RNA-dependent RNA polymerase (RdRp) of RNA viruses (2). It was discovered while screening chemical library for ~~anti-viral~~ antiviral effects on the influenza virus by Toyama Chemical Co., Ltd (2). Favipiravir is effective toward various types and subtypes of influenza viruses (3). It also exhibits antiviral properties against other RNA viruses, including filoviruses, arenaviruses, and bunyaviruses (4). Because of its distinct antiviral properties, favipiravir has the potential to be a promising treatment for untreatable RNA viral infections (4). It has shown efficacy against many RNA viruses, including Ebola virus, Lassa virus, norovirus and enterovirus *(5–8)*.

Additionally, favipiravir is approved for emergency use to treat COVID-19 disease (9–14). The SARS-CoV-2 virus is responsible for COVID-19 disease, which infects human instantaneously (15). This evolution from influenza treatment to a key therapeutic candidate for emerging RNA virus outbreaks exemplifies favipiravir's unique broad-spectrum antiviral profile and therapeutic promise. But like many drugs, it has some limitations, such as low bioavailability and side effects, which reduce its therapeutic effectiveness. Some reported side effects include diarrhea, elevated transaminase levels, hyperuricemia, nausea, thrombocytopenia and neutropenia (16, 17). To overcome these challenges advanced drug delivery systems employing nanomaterials have been developed. Drug delivery systems have attracted significant attention due to their ability to deliver precise amounts of drugs at specific target sites in the human body (18), minimizing the risk of side effects so drug delivery nanocarriers are essential for enhancing drug bioavailability and therapeutic effectiveness in treating specific viruses and diseases.

Various nanoparticles including gold nanoparticles, liposomes, silicone nanoparticles, carbon nanotubes, natural and synthetic polymers, as well as magnetic nanoparticles, have recently been employed for drug delivery applications (19, 20). Carbon materials such as fullerenes, carbon nanotubes, and graphene, in addition to their modified forms, have gained important attention recently as useful tools in medicine, especially as carriers that work for drug delivery (21, 22). Since their nanoscale dimensions permit them to penetrate biological membranes more easily, they improve drug transport along with help drugs stay in the body longer. Since carbon nanomaterials do have distinctive structural and physicochemical features, they stand out as superb platforms that allow for the development of advanced drug delivery systems (23, 24). These candidates are strong to make therapeutic delivery approaches that are more effective and efficient (25, 26). In addition, their performance is strongly influenced by the nature of intermolecular

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interactions governing drug adsorption and release. Intermolecular interactions play a fundamental role in determining the efficiency of drug-nanocarrier systems. The adsorption, stability, and release behavior of drug molecules are mainly governed by non-covalent interactions such as hydrogen bonding, π–π stacking, van de Waals forces, and electrostatic interactions. These interactions are particularly important in carbon-based nanocarriers, where π-conjugated surfaces enable strong yet reversible binding with aromatic or heterocyclic drug molecules. A balanced interaction strength is essential to ensure efficient drug loading while allowing controlled release at the target site. Therefore, understanding these intermolecular forces at the molecular level is crucial for designing effective nanocarriers with optimized therapeutic performance (27).

Graphdiyne (GDY) is a novel two-dimensional carbon-based nanomaterial with unique structural and physicochemical properties, including high surface area, tunable electrical conductivity, uniform pores, strong π-conjugation, high thermal stability, thermal conductivity, excellent biocompatibility, and low toxicity, which make it suitable for biomedical applications (28–30). Compared with other carbon nanomaterials such as graphene, carbon nanotubes, and fullerenes, graphdiyne (GDY) possesses uniformly distributed intrinsic pores and a mixed sp/sp² hybridized carbon network, which provide additional active adsorption sites and facilitate stronger interactions with drug molecules (31). These unique structural characteristics make GDY a particularly promising nanocarrier for enhanced drug loading and controlled drug release. Graphdiyne (GDY) has been widely investigated in biosensing, cancer therapy**,** drug delivery, wound healing, radiation protection, and tissue engineering, among these applications, its role in drug delivery is significant as graphdiyne (GDY) can load therapeutic molecules through π–π stacking and electrostatic interactions and release them in a controlled and targeted way (28–30, 32–35). These qualities enable graphdiyne (GDY) to improve the solubility, stability, and controlled release of drug at the target site, and enhancing therapeutic efficacy while decreasing toxicity, Studies show graphdiyne (GDY) nanosheets can effectively deliver drugs such as cisplatin drug (36) **,** Flutamide drug (37), Imuran, Pentasa and Hyoscyamine drugs (38), Temozolomide drug (39), Hydroxyurea and 5-Fluorouracil drugs (40). Additional functional improvements of graphdiyne (GDY) as a drug carrier can be achieved by decorating its nanosheet with transition metals. Metal decoration provides new active sites and modifies the electronic structure of the carbon nanosheets, which in turn enhances drug loading efficiency, targeting specificity, and therapeutic efficacy. The modified electronic properties enhance the

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adsorption energies and molecular interactions between graphdiyne and drug molecule, resulting in a more stable drug-nanocarrier complex and enabling controlled drug release(41–48).

Despite extensive DFT studies on Favipiravir adsorption over carbon nanomaterials such as graphene (49), carbon nanotubes (50), and fullerenes (51), the interaction of favipiravir with graphdiyne (GDY), particularly in its metal decorated forms, remains largely unexplored. In this study, a computational method density functional theory (DFT) was employed to investigate the adsorption of favipiravir on pristine and metal-decorated graphdiyne (GDY) as a carrier for the antiviral drug favipiravir. For decoration we use transition metals nickel (Ni), copper (Cu) and titanium (Ti). The choice of the specified transition metals for decoration on graphdiyne (GDY) due to their affinity to form stable complexes with organic compounds (52), as well as their ability to modulate their distinct electronic characteristics and interaction behavior with carbon-based systems. Ni exhibits strong d-π hybridization leading to stable binding and significant electronic coupling. Cu provides moderate interaction strength with relatively electronic conductivity, while Ti offers high chemical reactivity and significant charge transfer, allowing a systematic evaluation of metal-dependent effect on adsorption behavior (53). In addition, these transition metals have been widely employed in previous theoretical studies on carbon-based nanomaterials, where they were shown to significantly influence adsorption properties and electronic structure modulation(54–57). Recent studies have revealed that density functional theory (DFT) provides valuable insights into molecular structure and their interactions (58). Density functional theory (DFT) has been widely used to explore drug-nanocarrier interactions, particularly through the calculation of adsorption energies and analysis of molecular interactions (59–63). These parameters offer important details about the strength and stability of drug binding, which support the design of more effective drug delivery systems. Researchers can refine the molecular structure of nanocarriers to enhance their targeting capabilities and therapeutic efficacy in various biomedical applications. This can lead to improved treatment outcomes and more effective delivery of drugs (64, 65). To enhance the understanding of adsorption processes, additional analyses were performed. Molecular orbital analysis, including (HOMO-LUMO) and corresponding energy gaps, was employed to evaluate the electronic reactivity and charge transfer capability upon adsorption. Natural bond orbital (NBO) analysis was performed to determine donor-acceptor interactions and intramolecular charge delocalization, providing a detailed description of the stabilization arising from orbital interactions. Furthermore, density of states (DOS) analysis was used to investigate the changes in the electronic structure of the system. Infrared (IR) spectroscopy simulations

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were conducted to identify vibrational characteristics associated with adsorption. Additionally reduced density gradient and non-covalent interaction (RDG-NCI) analyses were performed to visualize and distinguish weak intermolecular forces, such as van der Waals forces and hydrogen bonding. Thermodynamic parameters were evaluated to assess the stability of adsorption process. Moreover, work function was performed to evaluate the influence of adsorption on the surface electronic properties. The goal of our research is to explore the properties of geometric parameters, adsorption energies, and various electrical and geometric characteristics using DFT.

2. *Computational methods*

Density functional theory (DFT) used to investigate the adsorption behavior of favipiravir drug on pristine and metal-decorated graphdiyne surfaces in gas and water phases. For optimization of geometry and analysis of electronic properties and energy computations along with the associated computations the B3LYP exchange–correlation functional was used together with the 6-31G(d) basis set (66), as implemented in the Gaussian 09 software package (67). This level of theory was selected due to its well-established balance between accuracy and computational efficiency and have been successfully used for the treatment of similar chemical systems. Visualization and analysis of optimized geometries were performed using Gauss View 6.0. However, unrestricted calculations were employed for the Cu complexes, as the presence of unpaired electrons. DFT was selected for this study because it provides an excellent balance between computational efficiency and accuracy in describing the electronic structure of complex systems. DFT offers reliable insights into adsorption energies, electronic distributions, and intermolecular interactions at a reasonable computational cost. This makes it particularly suitable for investigating drug adsorption on nanomaterial surfaces(68–80). The electronic properties were performed to evaluate the stability of the FAV–GDY complexes, including highest molecular orbital (HOMO) and lowest unoccupied molecular orbital (LUMO). Density of states (DOS), infrared (IR) spectra were analyzed using Gauss Sum. Reduced density gradient–noncovalent interaction (RDG–NCI) analysis to explore noncovalent interactions. Frequency calculations were performed to explore intermolecular interactions and to obtain the thermodynamic parameters associated with the adsorption process, including enthalpy change (ΔH), Gibbs free energy change (ΔG) and entropy (ΔS). Frequency calculations have been done on all chemical systems at the same level of theory for the optimization. All optimized structures were confirmed to correspond to true global minimum on the potential energy surface, as verified by vibrational frequency calculations showing no imaginary frequencies, indicating that the structures were

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actually minima. Through the analysis of adsorption energies and molecular descriptors, these calculations offered valuable insight into adsorption mechanisms.

The adsorption energy (E_{ad}) of the favipiravir molecule on pristine graphdiyne surface is calculated, for the selected configurations, through the following equation:

$$E_{ads} = E_{GDY/FAV} - (E_{GDY} + E_{FAV}) \quad (1)$$

Here, E_{GDY} is the total energy of graphdiyne, E_{FAV} is the total energy of favipiravir, and $E_{FAV/GDY}$ is the total energy of favipiravir adsorbed on graphdiyne surface, where negative values indicate favorable adsorption.

To improve the accuracy of the calculated interaction energies, basis set superposition error (BSSE) was corrected using the counterpoise method, as defined by the following equation:

$$E_{ads} = E_{GDY/FAV} - (E_{GDY} + E_{FAV}) + E_{BSSE} \quad (2)$$

The energy gap between HOMO and LUMO (Eg) and Fermi level energy are defined as follows:

$$E_{gap} = E_{LUMO} - E_{HOMO} \quad (3)$$

$$E_F = \frac{E_{LUMO}+E_{HOMO}}{2} \quad (4)$$

Where, E_{LUMO} and E_{HOMO} are energy of highest occupied molecular orbital (HOMO) and lowest unoccupied molecular orbital (LUMO).

3. *Result and discussion*

3.1. *Different optimized adsorption geometries of Favipiravir on GDY surface*

The interaction of favipiravir with the GDY nanosheet was explored through the optimization of different orientations. This approach enabled the identification of the most stable adsorption geometry and provided clearer insight into how the drug interacts with the GDY surface. Five different orientations of the drug molecule were investigated and optimized, as shown in **Fig. 1.** In Tilted-OH orientation (**Fig. 1a**), the hydroxyl group directed closer to the GDY surface, with interaction distances of 2.53 Å between hydrogen and carbon atom and 3.35 Å between oxygen atom and carbon. This arrangement allows simultaneous π–π stacking and dipole-dipole interactions. In Vertical-OH (**Fig. 1b**) hydroxyl remains the main contact site with the GDY surface, larger interaction distances observed, 2.84 Å between hydrogen and carbon

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atom and 3.40 Å between oxygen atom and carbon, reflecting a weaker adsorption efficiency. In NH_2/F-oriented (**Fig. 1c**), the drug molecule oriented perpendicularly to the GDY surface, the NH_2 and F groups are directed toward the surface, enabling additional polar interactions with the GDY surface. The interaction distances were 2.90 Å between the hydrogen of the NH_2 group and the GDY surface, and 3.26 Å between the fluorine atom and the GDY surface. In Carbonyl-side (**Fig. 1e**), the FAV shows adsorption along the edge of the GDY surface, where contact is established mainly through side interactions, this orientation supported mainly by weak van der Waals forces, with interaction distances of 2.34 Å between the carbonyl group and the hydrogen of the GDY nanosheet, and 3.49 Å between the nitrogen atom of the pyrazine ring and the carbon of the GDY nanosheet. The contact is less stable and less uniform compared to the other orientations. In Hydroxyl-side (**Fig. 1d**) the hydroxyl group is directed toward the edge of the GDY nanosheet creating hydrogen bond and dipole-dipole interactions, in addition to van der Waals interactions, with interaction distances of 2.45 Å between the oxygen atom and the hydrogen of the GDY nanosheet, and 2.51 Å between the hydrogen atom and the carbon of the GDY nanosheet. The hydroxyl-OH configuration was found to be the most stable orientation in Vacuum with adsorption energy − 0.26 eV, and shorter interaction distances compared to the other adsorption modes.

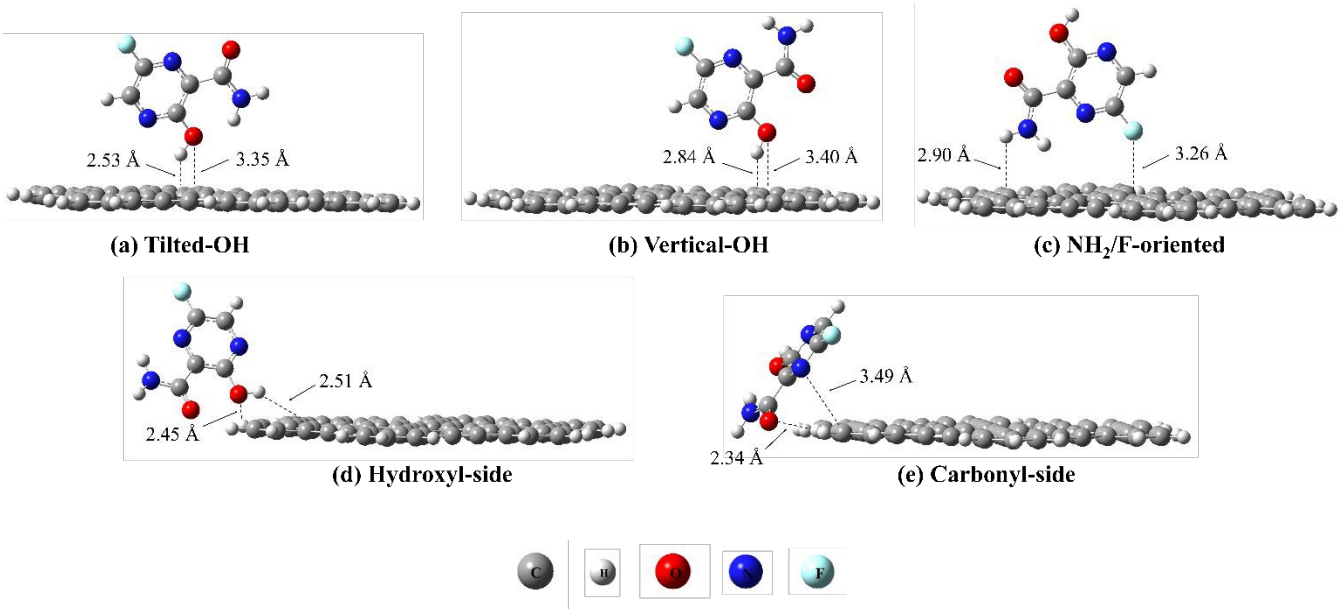

Figure 1. Optimized geometries of FAV drug on GDY nanosheet in five orientations, highlighting variations in molecular alignment and interaction sites responsible for adsorption behavior (in gas phase).

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To further enhance this interaction and explore possible improvements in carrier efficiency, metal decoration was subsequently introduced to the GDY framework. The decorated models were optimized using the same computational protocol, allowing a direct and reliable comparison with the pristine systems. This stepwise strategy not only clarifies how surface functionalization modifies the adsorption landscape but also demonstrates the extent to which decoration can strengthen the drug carrier interaction.

3.2. *Optimized geometries of FAV on decorated GDY*

To explore the impact of the decorated systems on the binding of our drug, the interaction between the drug and each decorated model has been carefully modeled. Two different orientations for each metal-decorated GDY have been obtained and thus a total of six models were identified. The optimized configurations of favipiravir on Ni, Cu, and Ti-decorated GDY are presented in **Fig. 2**, where two adsorption orientations were obtained for each metal. These geometries reveal how metal anchoring sites restructure the interaction landscape and significantly modify the binding characteristics of the drug. The interaction distances and corresponding adsorption energies highlight how the drug adopts different geometries to optimize binding. Across all systems, the first orientation consistently exhibits a more favorable adsorption energy, indicating a stronger and more stable interaction compared to the second orientation.

For Ni-decorated GDY, the adsorption of favipiravir is stabilized through direct coordination between the metal center and heteroatoms of the drug molecule. In the first orientation the Ni atom interacts with the nitrogen and oxygen atoms of favipiravir, forming Ni-O and Ni-N coordination bonds with interaction distances of 1.89 Å and 1.92 Å, respectively, these short interaction distances indicate strong metal-ligand coupling and result in a highly stable adsorption configuration. In the second orientation the Ni atom forms Ni-O and Ni-N coordination bonds with distances of 1.91 Å and 2.25 Å, respectively, remain within the range of strong surface coordination, the slightly elongated second distance corresponds to a less negative adsorption energy value.

In case of Cu-decorated GDY surface, the adsorption mechanism is similarly governed by coordination between Cu atom and heteroatoms of favipiravir. In the first orientation, the drug molecule establishes Cu-O and Cu-N interactions with distances of 1.98 Å and 2.01 Å, respectively. In the second orientation, the coordination framework remains intact, characterized by Cu-O and Cu-N coordination with interaction distances of 1.90 Å and 2.78 Å, respectively.

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For the Ti-decorated GDY, the adsorption of favipiravir is dominated by coordination interactions involving oxygen and nitrogen atoms. The first orientation exhibits Ti-O and Ti-N interactions with distances of 1.95 Å and 2.26 Å, respectively. In second orientation the interaction is characterized by Ti-O and Ti-N coordination with distances of 2.16 Å and 2.01 Å respectively. The adsorption behavior of favipiravir on the decorated GDY surfaces reveals a clear enhancement in interaction strength when metal decoration is introduced. All decorated systems exhibit markedly stronger adsorption energies compared to pristine GDY sheet, indicating that surface modification creates highly active binding centers capable of forming stronger drug-carrier interactions. The decorated structures show strong interactions, confirming the effectiveness of the modification strategy.

Among the evaluated systems, the FAV(N/O)-GDYNi, FAV(N/O)-GDYCu and FAV(N/O)-GDYTi complexes **Fig. 2 (a-c**), demonstrate the highest binding affinity toward favipiravir, indicating a substantial stabilization of favipiravir on the decorated nanosheet. These findings demonstrate that the metal decoration alters the electronic structure and reactivity of GDY, creating coordination sites capable of interacting with the oxygen and nitrogen atoms of favipiravir. The resulting interactions are considerably stronger than those observed on the pristine GDY surface, which supports the idea that decoration not only enhances binding affinity but also offers tunability depending on the selected metal.

Overall, the decorated GDY systems exhibit the essential characteristics of an efficient nanocarrier. The significant increase in adsorption energy relative to the pristine sheet underscores the promising potential of decorated-GDY as a highly effective nanocarrier for favipiravir, offering improved stability and stronger binding essential for efficient drug delivery applications.

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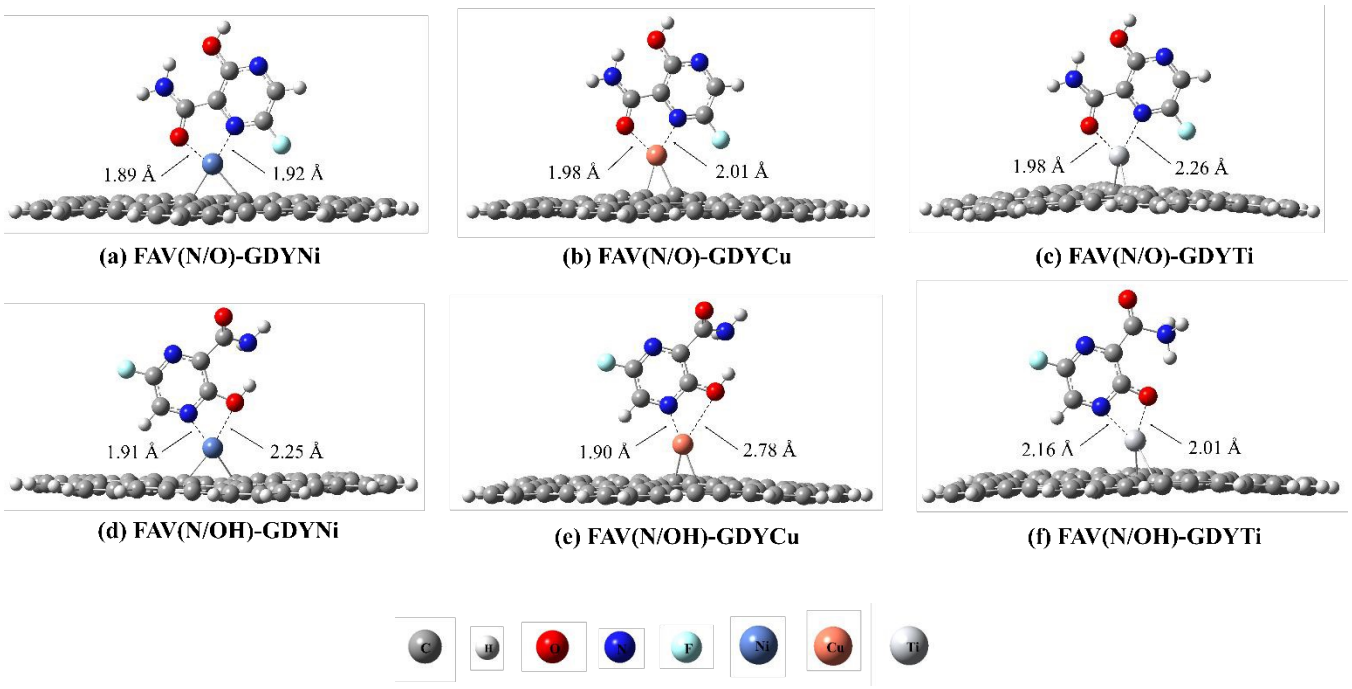

Figure 2. Optimized geometries for different adsorption orientations of FAV on decorated-GDY surfaces: Ni-GDY, Cu-GDY and Ti-GDY, with their interaction distances in gas phase.

3.3. *Frontier molecular orbital analysis*

The frontier molecular orbitals (FMOs), namely the highest occupied molecular orbital (HOMO) and the lowest unoccupied molecular orbital (LUMO), offer a detailed understanding of the electronic distribution and interaction mechanism between favipiravir and the GDY nanosheet. As shown in **Fig. 3 and 4**, the HOMO electron density is mainly localized on the favipiravir molecule, particularly around the oxygen and nitrogen sites, while the LUMO density is predominantly distributed over the GDY surface. The observed spatial separation of HOMO and LUMO orbitals reflect a pronounced donor-acceptor character within the FAV-GDY system, in which the favipiravir acts as the electron donor and GDY serves as the electron acceptor. For Cu complexes, both α- and β-spin configuration were analyzed, however, only the α-spin HOMO and LUMO are presented, as it demonstrates more pronounced variation in the electronic structure, it exhibits larger band gap values compared to β-spin configuration (see **Table 5**). The electron transfer tendency is further supported by the alignment of molecular orbitals, the localization of the LUMO on GDY nanosheet facilitates charge transfer from the drug molecule toward the nanosheet upon adsorption. This implies that the adsorption process involves charge transfer through non-covalent interactions, enhancing the stability of the complex. The delocalization of the LUMO orbitals across the

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π-system of GDY indicates its strong capacity to accept electrons, while the confinement of the HOMO on favipiravir highlights its donor character arising from the heteroatoms and functional groups such as NH_2, OH and C=O. These findings provide a deep understanding of the electronic features governing favipiravir adsorption on GDY and support its potential as adsorption material.

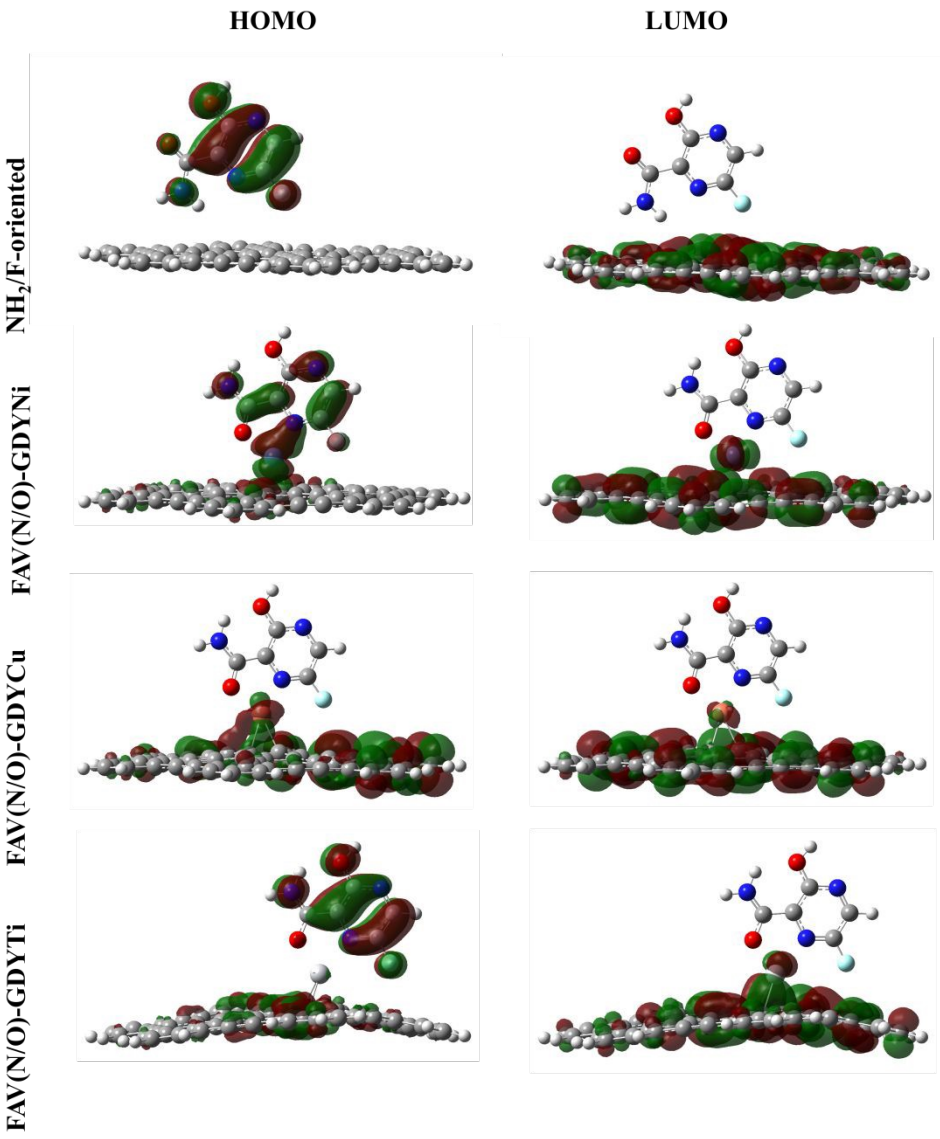

Figure 3. The distributions of molecular orbitals HOMO and LUMO of NH_2/F-oriented, FAV(N/O)-GDYNi, FAV(N/O)-GDYCu and FAV(N/O)-GDYTi complexes.

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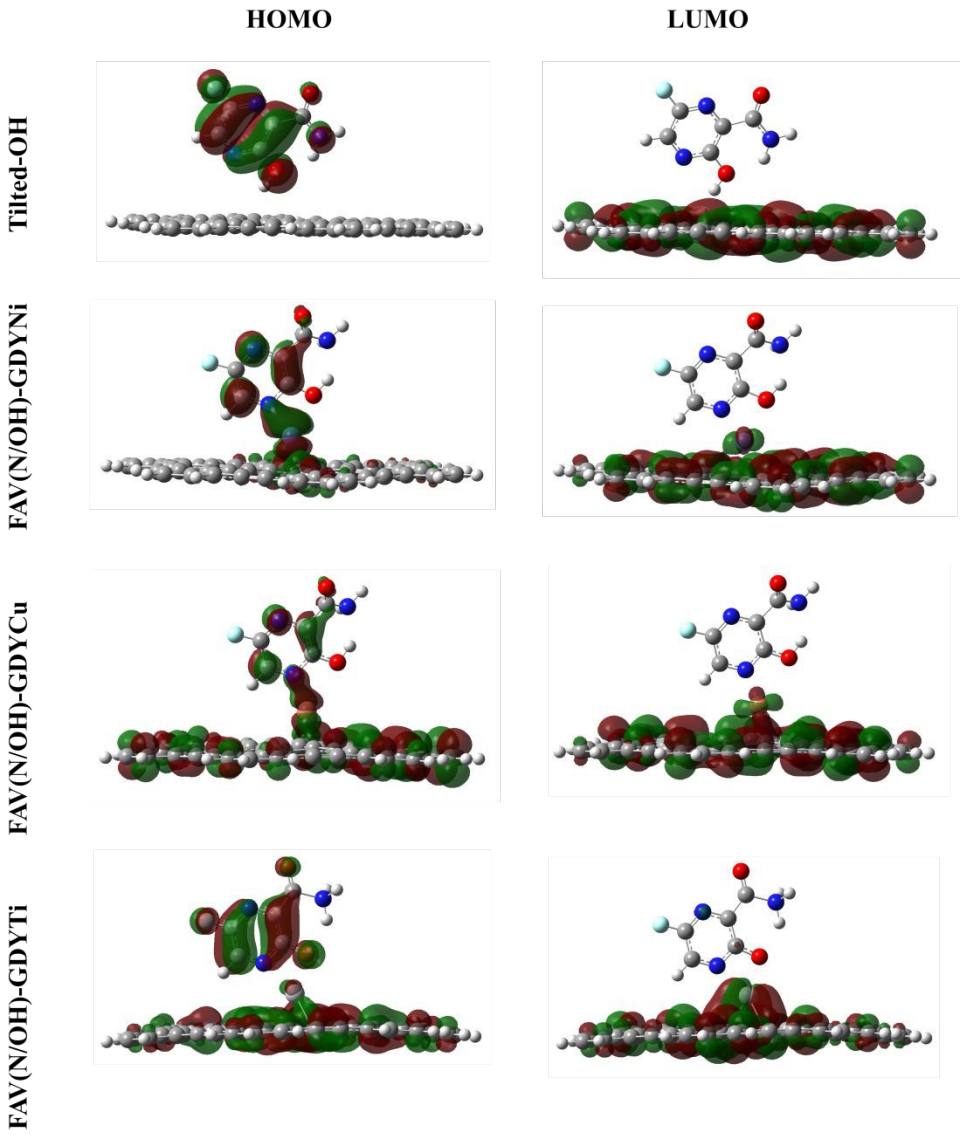

Figure 4. The distributions of molecular orbitals HOMO and LUMO of the Tilted-OH, FAV(N/OH)-GDYNi, FAV(N/OH)-GDYCu and FAV(N/OH)-GDYTi complexes.

3.4. *Electronic and vibrational analysis (DOS and IR)*

The combined analysis of the Density of states (DOS) and infrared (IR) spectra provides comprehensive insight into both the electronic structure and vibrational behavior of molecular system. DOS analysis reveals detailed information about electronic distribution and orbital interactions (81–85). The DOS spectra illustrate how the molecule orbitals of the drug interact with those of the nanosheet, revealing the nature of the electronic coupling in the system. The DOS profile shows a continuous distribution of states

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around the fermi level, suggesting an effective overlap between the orbitals of the favipiravir and those of the nanosheet. The infrared (IR) spectra were also analyzed to identify vibrational changes associated with adsorption and to confirm stability of the optimizes geometries. **Fig. 5, 6.** presented the simulated spectra, where distinct vibrational modes corresponding to characteristic functional groups of favipiravir, such as C=O, N–H, and C–F stretching vibrations. When studied together, DOS and IR analyses reveal both the electronic changes and structural stability resulting from molecular adsorption, which are critical in evaluating the suitability of nanomaterials for drug delivery.

The electronic properties were further analyzed using energy gaps (Eg), and Fermi level (EF), which provide detailed insight into the interaction between favipiravir (FAV) and GDY-based nanostructures, as illustrated in Table 5. For the pristine GDY–FAV configurations, the calculated energy gaps fall within a narrow range of 2.498–2.611 eV, while the Fermi level varies slightly between −4.128 and −3.844 eV. These minor variations indicate weak interaction and limited charge transfer between FAV and the pristine GDY surface, with negligible perturbation in the electronic structure.

In contrast, the metal-decorated GDY–FAV complexes exhibit a pronounced modification in electronic properties. For FAV(N/O)-GDYNi, the energy gap decreases significantly to 0.813 eV with E_F at −9.422 eV. The FAV(N/O)-GDYCu systems show moderate E_g values of 1.531 eV and 1.019 eV, with corresponding E_F values of −8.832 eV and −9.059 eV, depending on the adsorption configuration. Notably, the FAV(N/O)-GDYTi complex exhibits the lowest energy gap (0.328 eV) with a deeply shifted Fermi level (−14.248 eV), indicating strong electronic interaction and superior charge transfer capability. A similar trend is observed for FAV(N/OH)-GDYNi, which shows Eg of 0.724 eV and E_F of −9.406 eV. For FAV(N/OH)-GDYCu, the energy gap varies significantly depending on the configuration, ranging from 2.159 to 0.596 eV, with E_F values between −8.837 and −9.572 eV, indicating configuration-dependent interaction strength. Meanwhile, FAV(N/OH)-GDYTi maintains a low Eg of 0.344 eV with EF at −14.091 eV, confirming the stability of its enhanced electronic behavior.

Overall, the pristine GDY–FAV systems exhibit weak interaction characterized by minor changes in Eg and EF, whereas metal-decorated systems especially Ti-based complexes show a substantial reduction in energy gap and significant shifts in Fermi level. This indicates strong orbital hybridization and efficient charge transfer upon adsorption, which is expected to induce notable changes in DOS and enhance the suitability of these systems.

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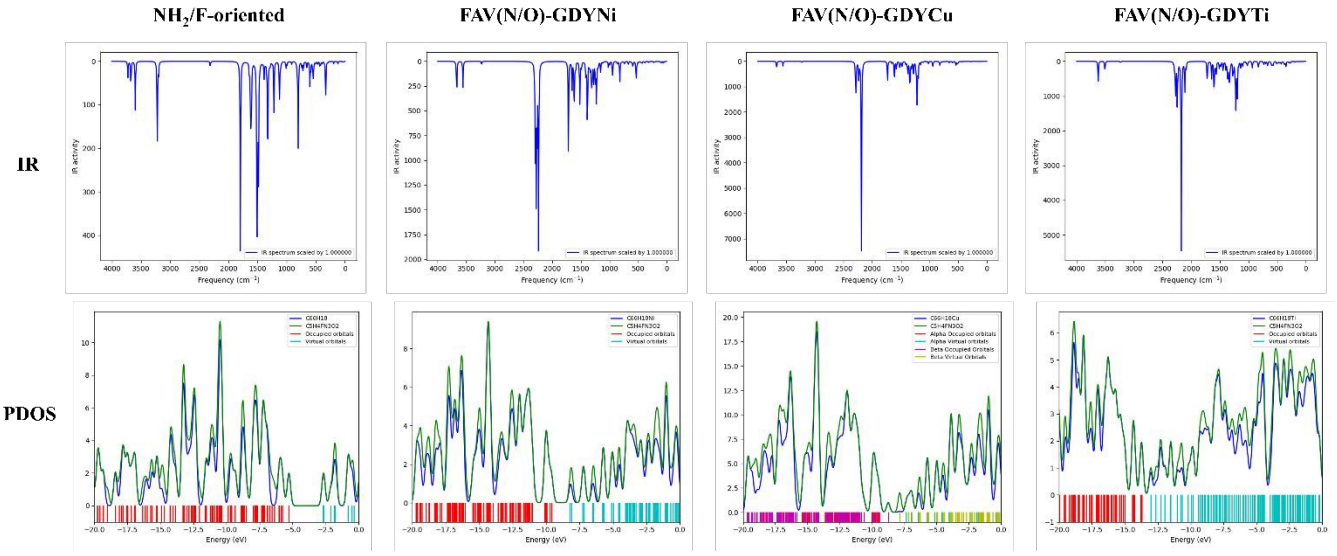

Figure 5. IR spectra and PDOS for NH_2/F-oriented, FAV(N/O)-GDYNi, FAV(N/O)-GDYCu and FAV(N/O)-GDYTi complexes.

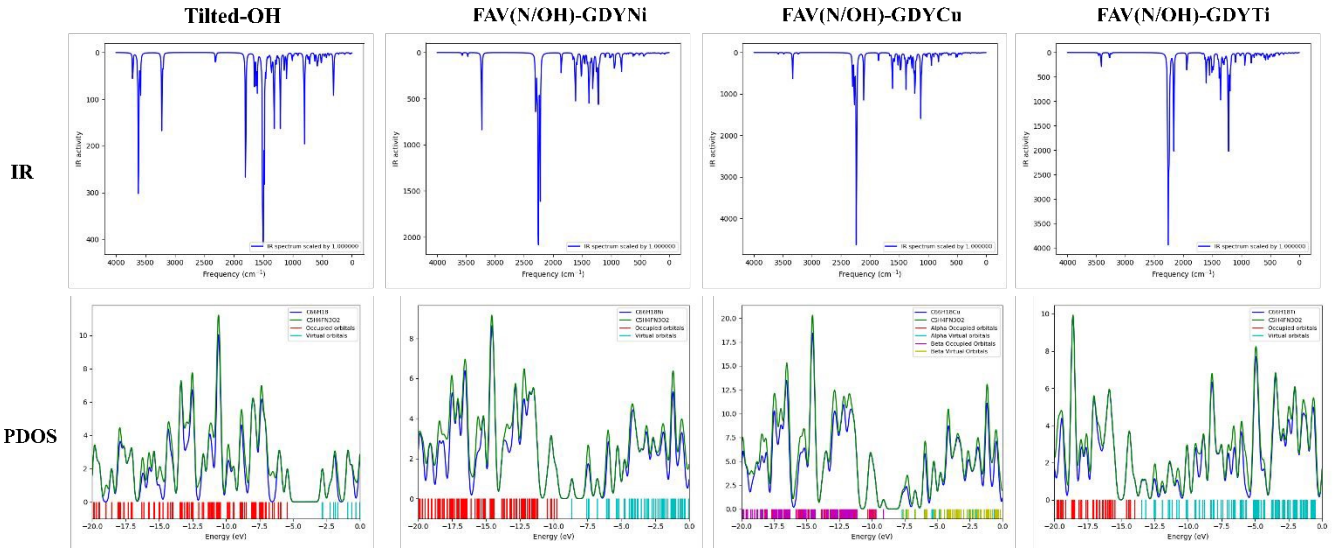

Figure 6. IR spectra and PDOS for tilted-OH, FAV(N/OH)-GDYNi, FAV(N/OH)-GDYCu and FAV(N/OH)-GDYTi complexes.

3.5. *Non-covalent interactions (NCI)*

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The non-covalent interaction (NCI) analysis combined with the reduced density gradient (RDG) plots was employed to elucidate the nature and strength of interactions between the favipiravir and the GDY surface. These analyses provide a clear visualization of the weak intermolecular forces responsible for stabilizing the drug-GDY complexes, such as van der Waals interaction, hydrogen bonding, and steric repulsions (86–91). It is important to note that NCI analysis is inherently limited to non-covalent interactions and does not account for covalent bonding or charge-transfer effects. Accordingly, for pristine GDY, where adsorption is dominated by weak interactions, NCI provides an appropriate description. In contrast, for metal-decorated GDY systems, additional covalent contributions arising from metal–drug interactions are expected but are not captured by NCI analysis. The NCI isosurface of the optimized configurations of FAV on pristine and decorated-GDY clearly demonstrates regions of intermolecular interaction. As shown in **Fig.7,8**, the visualization of NCI reveals predominant green region at the drug-surface interface, signifying that van der Waals forces are the main contributors to the stabilization of the drug on GDY nanosheet. This observation reflects the dominant non-covalent component of interaction and should not be interpreted as evidence of purely physisorptive binding, particularly in the decorated systems. The RDG scatter plots allow the identification of different types of non-covalent interactions through the sign (λ_2) ρ descriptors and color mapping. The blue regions suggest the presence of attractive hydrogen bonding between the oxygen or nitrogen atoms of FAV and the carbon atoms on GDY, while red areas correspond to repulsive steric effect arising from short interatomic distances. The distribution and intensity of these colored isosurface reveal that all configurations are stabilized mainly through non-covalent interactions, typically associated with hydrogen bond, while red areas correspond to repulsion steric effect. The RDG is a function that mainly depends on the electron density (ρ) and the second derivative of the electron ($\Delta^2\rho$), which is calculated using this equation:

$$\mathrm{RDG}\,(r) = \frac{1}{2(\,3\pi^2)^{1/3}} \frac{|\nabla\rho\,(r)|}{\rho(r)^{4/3}} \qquad (5)$$

where ρ represents the electron density and $\Delta\rho$ is its gradient at a given point. Low values of the reduced density gradient in regions of low electron density indicate the presence of non-covalent interactions. To differentiate between attractive, van der Waals, and repulsive interactions, the sign of the second eigenvalue of the electron density Hessian (λ_2) multiplied by ρ was considered. In the resulting RDG-NCI plots and isosurfaces, blue regions corresponded to strong attractive interactions such as hydrogen bonding, green regions indicated weak van der Waals interactions, and red regions represented steric repulsion. The analysis revealed that the complexes were stabilized predominantly by van der Waals

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forces, with hydrogen bonding contributing to additional stabilization, while steric repulsions were minimal. This visualization not only confirms the presence and type of non-covalent interactions but also provides insights into the spatial distribution of these interactions, which is critical for understanding the structural and energetic behavior of the complexes. However, when considered alongside the large binding energies, NBO stabilization energies, and work function shifts, it is evident that stronger interactions particularly in metal-decorated systems also involve partial covalent character and charge transfer. Therefore, adsorption on pristine GDY can be described as physisorption, whereas adsorption on metal-decorated GDY follows a mixed mechanism involving both non-covalent and chemisorption contributions.

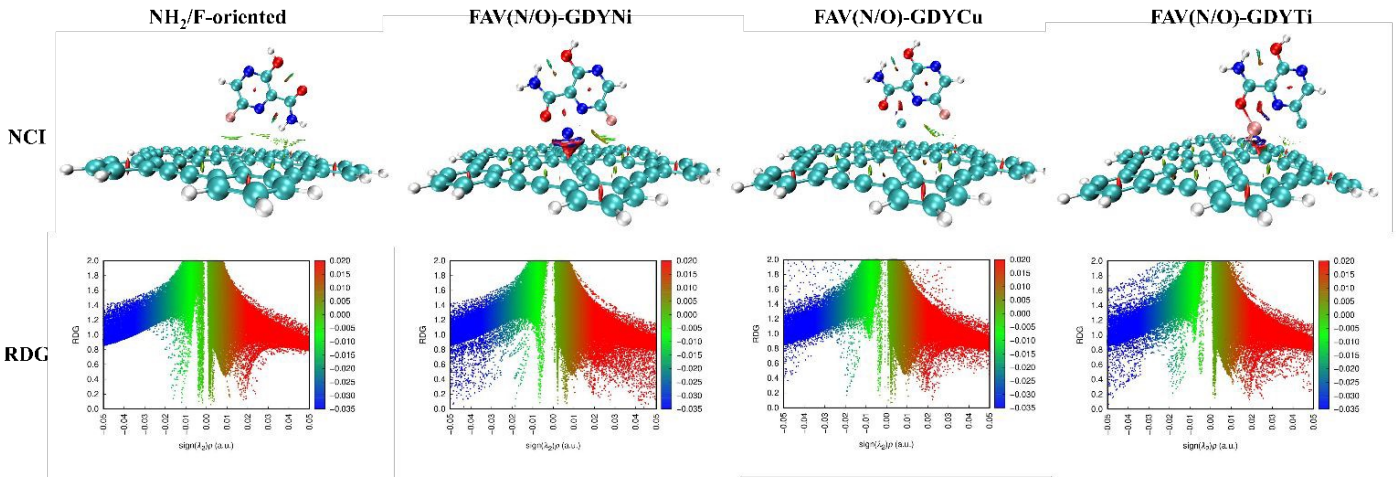

Figure 7. Non-Covalent Interaction (NCI), Reduced Density Gradients (RDG) for NH_2/F-oriented, FAV(N/O)-GDYNi, FAV(N/O)-GDYCu and FAV(N/O)-GDYTi complexes.

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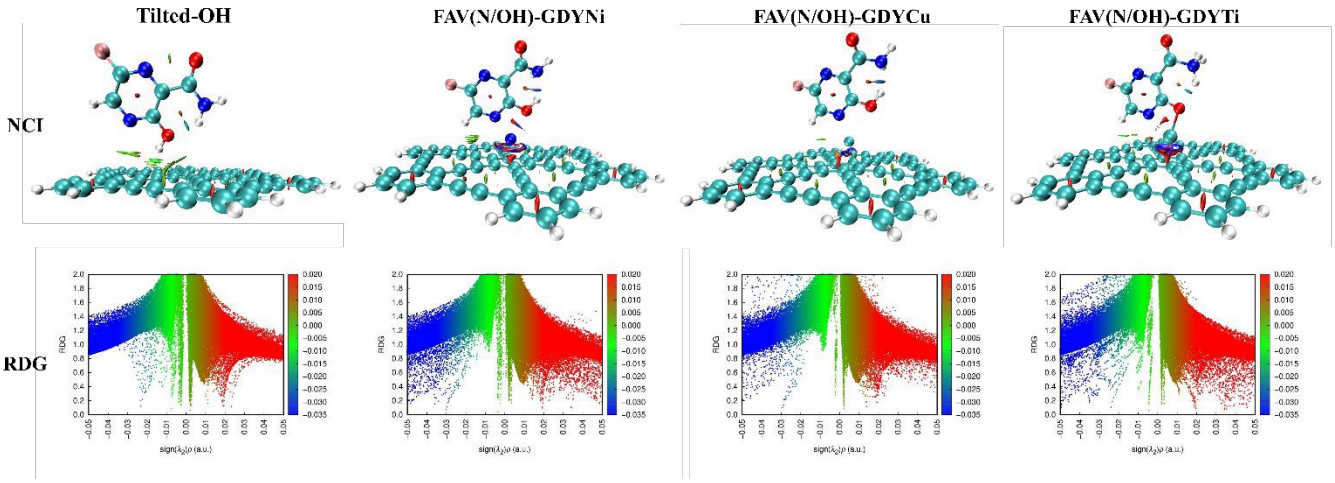

Figure 8. Non-Covalent Interaction (NCI), Reduced Density Gradients (RDG) for tilted-OH, FAV(N/OH)-GDYNi, FAV(N/OH)-GDYCu and FAV(N/OH)-GDYTi complexes.

3.6. *Thermodynamic parameters*

Thermodynamic analysis provides valuable insights into the energetic and spontaneous nature of the adsorption process between favipiravir and the graphdiyne surfaces. The calculated thermodynamic parameters, including Gibbs free energy (ΔG), enthalpy change (ΔH) and entropy change (ΔS), in both the gas and water phases are summarized in **Table 1**.

For pristine GDY surface, all examined orientations exhibit positive ΔG values, indicating that adsorption does not occur spontaneously whereas the ΔH values are slightly exothermic. This difference arises because the binding energy reflects only the electronic stabilization upon adsorption, whereas the Gibbs free energy also includes thermal and entropic contributions. For pristine GDY, these thermal effects outweigh the electronic stabilization, resulting in a positive ΔG value demonstrating a nonspontaneous process.

When the GDY surface is decorated with metal atoms, a notable enhancement in the thermodynamic behavior is observed. All decorated systems exhibit negative ΔG values, confirming the spontaneous nature of adsorption after modification. The significantly large negative ΔH values highlight the stronger and more stabilized interactions. The observed negative ΔS values for all complexes indicate a decrease in randomness upon adsorption.

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The enthalpy change (ΔH) and Gibbs free energy change (ΔG) were calculated at T = 298.15 K and P = 1 atm as follows:

$$\Delta G = G_{GDY/FAV} - (G_{GDY} + G_{FAV}) \quad (6)$$

$$\Delta H = H_{GDY/FAV} - (H_{GDY} + H_{FAV}) \quad (7)$$

$$\Delta S = \frac{\Delta H - \Delta G}{298.15} \quad (8)$$

Table 1. The calculated thermodynamic parameters, Gibbs free energy (ΔG), enthalpy (ΔH) and entropy (ΔS) of all studied complexes in (KJ/mol), FAV drug with pristine and metal decorated GDY nanosheets, in gas phase as well as in water, respectively.

Structure	Gas			Water		
	ΔG	**ΔS**	**ΔH**	**ΔG**	**ΔS**	**ΔH**
Tilted-OH	17.45	− 0.11	− 16.82	30.37	− 0.12	− 4.31
Vertical-OH	18.03	− 0.10	− 10.86	31.05	− 0.11	− 2.57
NH_2/F-oriented	26.47	− 0.12	− 8.62	44.54	− 0.17	− 4.93
Hydroxyl-side	16.90	− 0.12	− 19.47	31.35	− 0.12	− 4.79
Carbonyl-side	25.74	− 0.13	− 13.45	31.51	− 0.12	− 5.35
FAV(N/O)-GDYNi	− 366.72	− 0.15	− 411.15	− 224.23	− 0.14	− 267.20
FAV(N/O)-GDYCu	− 327.23	− 0.14	− 369.67	− 220.36	− 0.17	− 272.33
FAV(N/O)-GDYTi	− 519.41	− 0.17	− 570.17	− 189.69	− 0.18	− 244.75
FAV(N/OH)-GDYNi	− 214.39	− 0.15	− 259.53	− 89.51	− 0.18	− 142.08
FAV(N/OH)-GDYCu	− 218.12	− 0.13	− 256.77	− 130.57	− 0.17	− 181.54
FAV(N/OH)-GDYTi	− 489.12	− 0.19	− 544.30	− 38.47	− 0.19	− 95.42

3.7. *Solvent effect*

To clarify the adsorption behavior in the presence of a polar environment. The geometries of the FAV drug with GDY and metal decorated GDY complexes were reoptimized in the water medium by utilizing the Integral Equation Formalism Polarizable Continuum Model (IEFPCM) method, applying the hydride B3LYP functional along with the 6-31G(d) basis set. The structural response of all systems was examined

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upon introducing water as the solvent. As shown in **Fig. 9,10**, geometry optimizations showed that both pristine and metal-decorated GDY retain the same structural features with interatomic distances remaining close to those observed in the gas phase, confirming that the implicit solvent does not induce any noticeable geometric changes to the optimized models, An exception is observed for the FAV(N/OH)-GDYTi complex in gas phase **Fig. 2. (f)**, where an intramolecular hydrogen shift occurs, leading to the formation of a new bond between the hydrogen atom and nitrogen atom within the same molecule. Moreover, the effect of water as the solvent was evaluated through the calculated binding energies (see **Table 2**). For the pristine GDY, the interaction remains exothermic, as indicated by the negative binding energy obtained.

Table 2. The binding energies (E_{bind} in eV) of all studied complexes in gas and water phase and solvation energy.

System	E_{bind} (eV)		
	Without BSSE correction	With BSSE correction	Presence of solvent
Tilted-OH	− 0.23	− 0.12	− 0.095
Vertical-OH	− 0.16	− 0.06	− 0.079
NH_2/F-oriented	− 0.14	− 0.01	− 0.074
Hydroxyl-side	− 0.26	− 0.14	− 0.108
Carbonyl-side	− 0.20	− 0.07	− 0.113
FAV(N/O)-GDYNi	− 4.41	− 3.35	− 2.926
FAV(N/O)-GDYCu	− 3.96	− 2.88	− 2.928
FAV(N/O)-GDYTi	− 5.79	− 5.39	− 2.696
FAV(N/OH)-GDYNi	− 2.75	− 2.00	− 1.547
FAV(N/OH)-GDYCu	− 2.74	− 2.09	− 1.926
FAV(N/OH)-GDYTi	− 5.79	− 5.45	− 1.117

This confirms that the solvent does not suppress the intrinsic adsorption tendency of the pristine surface to interact with the adsorbate. However, once metal decoration is introduced, the binding energies become

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more negative across all decorated configurations, demonstrating a clear enhancement in interaction strength and more stabilized adsorption process.

The thermodynamic parameters obtained in water further support the improvement introduced by metal decoration, as detailed in **Table 1**. For pristine GDY, the adsorption process exhibits a positive Gibbs free energy, including that the reaction is not spontaneous, even though the enthalpy change is slightly exothermic. The negative entropy values are consistent with the expected reduction in molecular freedom upon adsorption. Upon decoration, a clear shift in the thermodynamic behavior is observed. Gibbs free energy becomes negative for all decorated systems, demonstrating that adsorption becomes spontaneous under the same solvated conditions. Enthalpy values become more exothermic, providing the dominant contribution that drives the overall feasibility of the reaction. While entropy remains negative, enhanced enthalpic stabilization outweighs this contribution, yielding a thermodynamically favorable process.

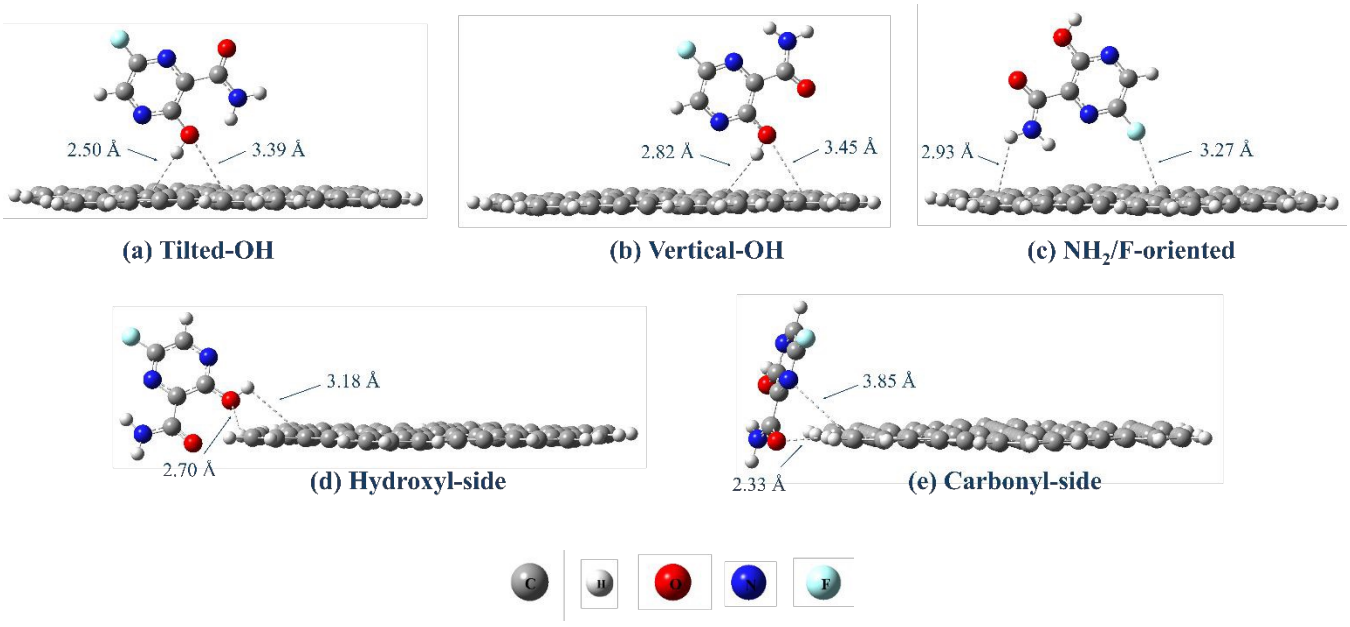

Figure. 9. Optimized geometries of FAV drug on GDY nanosheet in five orientations, with their interaction distances in solvent.

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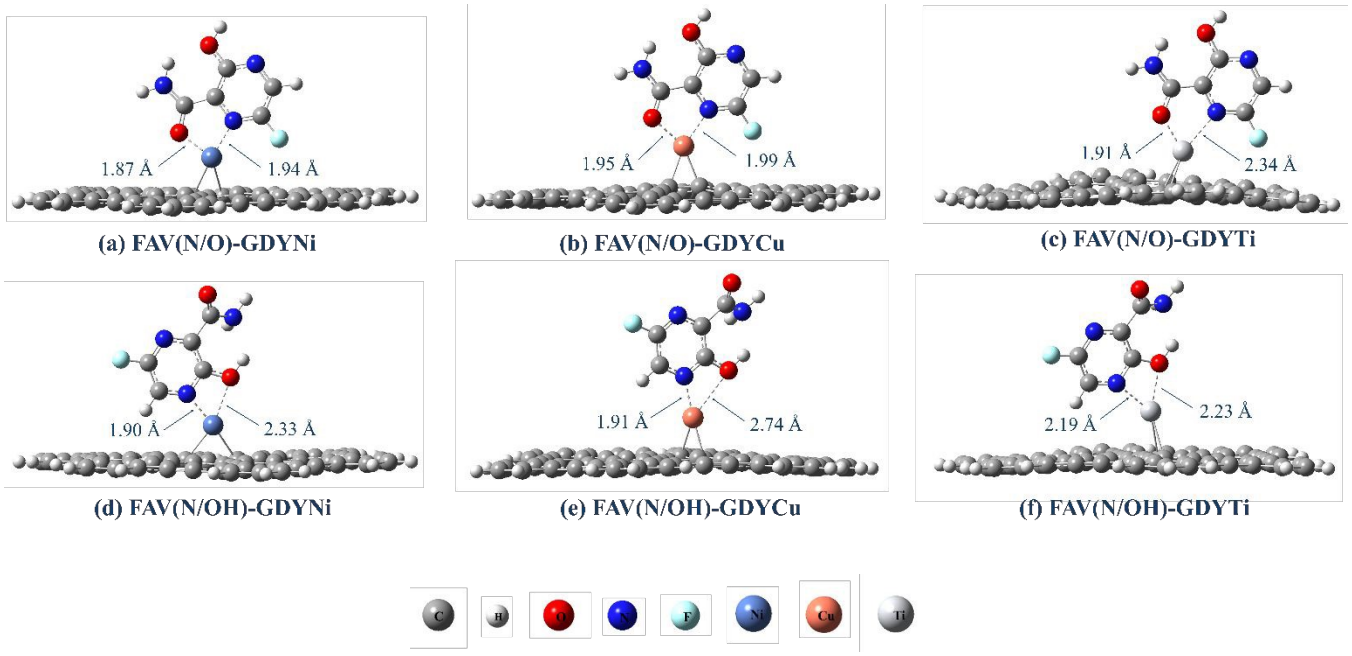

Figure 10. Optimized geometries for different adsorption orientations of FAV on decorated-GDY surfaces: Ni-GDY, Cu-GDY and Ti-GDY, with their interaction distances in solvent.

3.8. *Global indices parameters*

The set of global reactivity parameters presented in **Table 3**, including ionization potential (IP), electron affinity (EA), global hardness (η), softness (χ), chemical potential (μ), and electrophilicity index (ω) used to evaluate how the electronic structure of favipiravir response during adsorption. These descriptors help clarify the ability of molecules to donate or accept electrons (IP and EA), Its overall stability and resistance to electronic change (η and χ), and its tendency to attract electron density (μ and ω). Examining these parameters allows a clearer understanding of how the drug adjusts its electronic behavior in different orientations and when interacting with the modified GDY surface.

$$IP = -E_{HOMO} \quad (9)$$

$$EA = -E_{LUMO} \quad (10)$$

$$\eta = \frac{IP - EA}{2} \quad (11)$$

$$\chi = \frac{1}{\eta} \quad (12)$$

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$$\mu_{ch} = \frac{IP + EA}{2} \quad (13)$$

$$\omega = \frac{\mu^2}{2\eta} \quad (14)$$

Table 3. The calculated values of ionization potential (IP), electron affinity (EA), global hardness (η), softness (χ), chemical potential (μ), and electrophilicity index (ω) for FAV with pristine and metal decorated GDY nanosheets.

System	IP (eV)	EA (eV)	η (eV)	χ $(eV)^{-1}$	μ_{ch} (eV)	ω (eV)
Tilted-OH	5.378	2.879	1.249	0.800	-4.128	6.822
Vertical-OH	5.284	2.672	1.305	0.765	-3.978	6.061
NH_2/F-oriented	5.335	2.741	1.296	0.771	-4.038	6.286
Hydroxyl-side	5.149	2.539	1.304	0.766	-3.844	5.663
Carbonyl-side	5.182	2.605	1.288	0.775	-3.894	5.883
FAV(N/O)-GDYNi	9.829	9.015	0.406	2.45	-9.422	109.094
FAV(N/O)-GDYCu	9.598^{α}	8.066^{α}	0.765^{α}	1.305^{α}	-8.832^{α}	50.923^{α}
	9.568^{β}	8.549^{β}	0.509^{β}	1.962^{β}	-9.059^{β}	80.517^{β}
FAV(N/O)-GDYTi	14.412	14.083	0.164	6.079	-14.248	617.105
FAV(N/OH)-GDYNi	9.768	9.044	0.362	2.760	-9.406	122.107
FAV(N/OH)-GDYCu	9.917^{α}	7.757^{α}	1.079^{α}	0.925^{α}	8.837^{α}	36.159^{α}
	9.870^{β}	9.273^{β}	0.509^{β}	3.350^{β}	-9.572^{β}	153.48^{β}
FAV(N/OH)-GDYTi	14.264	13.919	0.172	5.801	-14.091	575.994

3.9. *Natural bond orbital*

The Natural Bond Orbital (NBO) analysis was performed to obtain a detailed description of the electronic interactions between favipiravir and GDY-nanosheet. This approach provides a quantitative description of electron delocalization through interactions between occupied Lewis-type orbitals (donors) and unoccupied non-Lewis's orbitals (acceptors). All NBO calculations were performed using density functional theory (DFT) at the B3LYP/6-31G(d) level of theory, which is well established for describing

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orbital interactions and electronic delocalization. The interaction parameters, including the second-order perturbation stabilization energy (E^2), the energy difference between donor and acceptor orbitals ($E_j - E_i$) , and the off-diagonal Fock matrix element F_{ij}, were evaluated to quantify the strength of donor-acceptor interactions. These parameters are summarized in **Table 4**.

The stabilization energy associated with each donor-acceptor interaction is estimated using second-order perturbation theory, expressed as:

$$E^{(2)} = q_i \frac{F_{ij}^2}{E_j - E_i} \qquad (15)$$

where q_i is the occupancy of the donor orbital, E_i and E_j are the diagonal elements (orbital energies) of the donor and acceptor NBOs, respectively, and F_{ij} represents the off-diagonal Fock matrix element. Larger $E^{(2)}$ values indicate stronger donor-acceptor interactions and enhanced electronic stabilization of the complex. Among all investigated complexes, the strongest stabilization interaction is observed between LP (2) O and LP*(2) Ti in the FAV(N/O)-GDYTi complex, with a remarkably high stabilization energy of 71.79 kcal/mol, indicating pronounced charge delocalization and strong orbital overlap. Additionally, significant stabilization energies are found between LP (1) N and LP*(6) Ni interaction in the FAV(N/O)-GDYNi complexes, reaching values up to 69.46 kcal/mol, further confirming strong donor-acceptor coupling.

Table 4. Natural bond orbital analysis for the donor and acceptor interactions, as well as the second-order perturbation stabilization energy ($E^{(2)}$, kcal/mol), correspond to the charge transfer of FAV drug with pristine and metal decorated GDY nanosheets in gas phase.

System	Donor NBO (i)	Acceptor NBO (j)	$E^{(2)}$ (kcal/mol)	$E_{(j)}$-$E_{(i)}$ a.u.	$F_{(i,j)}$ a.u.
Vertical OH	BD(3)C33- C57	BD*(1)C86 -O98	0.06	0.65	0.005
	BD(3)C22-C 32	LP*(1) H99	3.18	0.42	0.036
Tilted-OH	BD(3)C21-C55	BD*(1)C86-O98	0.05	0.67	0.005
	BD(3)C20-C 31	LP*(1) H99	0.83	0.41	0.018
NH_2/F-oriented	LP(1)N95	BD*(3)C21 - C55	0.08	0.36	0.005

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	BD(3)C21-C55	BD*(1)N95-H96	0.50	0.77	0.018
Hydroxyl-side	BD(2)C59-C 78	LP*(1)H99	2.41	0.40	0.031
	LP(1)O98	BD*(1)C37-H 42	1.14	1.11	0.032
Carbonyl-side	LP(1)N90	BD*(1)C37-H 42	0.12	0.88	0.010
	LP(1)O94	BD*(1)C37-H 42	1.30	1.20	0.035
FAV(N/O)-GDYNi	LP(1)N91	LP*(6)Ni85	62.94	0.59	0.174
	LP(2)O95	LP*(6)Ni85	48.06	0.60	0.153
FAV(N/O)-GDYCu	LP(1)N90	LP*(6)Cu100	23.40	0.49	0.138
	LP(2)O94	LP*(6) Cu100	18.51	0.47	0.120
FAV(N/O)-GDYTi	LP(1)N91	LP*(4) Ti85	21.70	0.59	0.104
	LP(2)O95	LP*(2)Ti85	71.79	0.54	0.175
FAV(N/OH)-GDYNi	LP(1)N88	LP*(5)Ni85	69.46	0.56	0.177
	LP(1)O99	LP*(6)Ni85	12.61	0.89	0.095
FAV(N/OH)-GDYCu	LP(1)N87	LP*(6)Cu100	34.41	0.47	0.162
	LP(1 O98	LP*(7)Cu100	2.78	0.75	0.058
FAV(N/OH)-GDYTi	LP(1)N87	LP*(1) Ti 100	40.12	0.37	0.110
	LP(2)O98	LP*(3)Ti100	51.15	0.43	0.132

3.10. *Work Function*

The work function (ϕ) was evaluated to examine the effect of favipiravir adsorption on the electronic surface properties of GDY. This parameter is directly related to the minimum energy required to remove

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an electron from the fermi level to the vacuum level and therefore provides useful insights into changes in surface electronic behavior upon interaction.

The work function was obtained using the relation:

$$\phi = V_{vac(\infty)} - E_F \qquad (16)$$

where $V_{vac(\infty)}$ is the electrostatic potential at the vacuum level and E_F is the Fermi level of the system. Since the vacuum potential is defined at a position for the surface where the potential approaches zero, the work function can be approximated as the negative of the fermi energy, i.e., $\phi = -E_F$. The calculated values listed in table 4, show that the different favipiravir orientations on pristine exhibit relatively close work function values. After adsorption on decorated GDY, noticeable shifts in the Fermi energy are observed, which are directly reflected in the calculated work function values. These variations indicate that the interaction between favipiravir and the modified GDY surface alters the energy required for electron emission. Such changes in work function suggest a redistribution of electronic charge at the interface, which is a typical consequence of adsorption induced electronic coupling.

Additionally, the percentage change in work function was evaluated using the following equation:

$$\%\Delta\phi = \frac{\phi_{complex} - \phi_{nanosheet}}{\phi_{nanosheet}} \times 100 \qquad (17)$$

Where $\phi_{Complex}$ and $\phi_{nanosheet}$ represent the work functions of the adsorbed system and the pristine surface, respectively. This parameter provides a quantitative measure of how adsorption affects the surface electronic properties.

Overall, the observed changes in work function confirm that favipiravir adsorption induces measurable electronic modification at the GDY surface. These results highlight the importance of work function analysis in understanding surface electronic response upon interaction.

Table 5. The calculated values of HOMO energies (E_{HOMO}), LUMO energies (E_{LUMO}), energy gap (E_g) in eV and fermi level E_F in gas phase.

System	E_{HOMO}	E_{LUMO}	Eg	E_F (eV)	ϕ (eV)	Δ ϕ (%)
Pristine	− 5.214	− 2.600	2.614	− 3.907	3.907	−

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DOI: 10.1039/D6NA00232C

GDYNi	− 9.941	− 9.414	0.526	− 9.678	9.678	−
GDYCu	− 9.965$^{\alpha}$	− 8.195$^{\alpha}$	1.770$^{\alpha}$	− 9.080$^{\alpha}$	9.080$^{\alpha}$	−
	− 9.893$^{\beta}$	− 9.337$^{\beta}$	0.555$^{\beta}$	− 9.615$^{\beta}$	9.615$^{\beta}$	
GDYTi	− 14.536	− 14.194	0.342	− 14.365	14.365	−
Tilted-OH	− 5.378	− 2.879	2.498	− 4.128	4.128	5.668
Vertical-OH	− 5.284	− 2.672	2.611	− 3.978	3.978	1.817
NH_2/F-oriented	− 5.335	− 2.741	2.593	− 4.038	4.038	3.342
Hydroxyl-side	− 5.149	− 2.539	2.609	− 3.844	3.844	− 1.615
Carbonyl-side	− 5.182	− 2.605	2.577	− 3.894	3.894	− 0.341
FAV(N/O)-GDYNi	− 9.829	− 9.015	0.813	− 9.422	9.422	− 2.638
FAV(N/O)-GDYCu	− 9.598$^{\alpha}$	− 8.066$^{\alpha}$	1.531$^{\alpha}$	− 8.832$^{\alpha}$	8.832$^{\alpha}$	− 2.728$^{\alpha}$
	− 9.568$^{\beta}$	− 8.549$^{\beta}$	1.019$^{\beta}$	− 9.059$^{\beta}$	9.059$^{\beta}$	− 5.782$^{\beta}$
FAV(N/O)-GDYTi	− 14.412	− 14.083	0.328	− 14.248	14.248	− 0.817
FAV(N/OH)-GDYNi	− 9.768	− 9.044	0.724	− 9.406	9.406	− 2.807
FAV(N/OH)-GDYCu	− 9.917$^{\alpha}$	− 7.757$^{\beta}$	2.159$^{\alpha}$	− 8.837$^{\alpha}$	8.837$^{\alpha}$	− 2.671$^{\alpha}$
	− 9.870$^{\alpha}$	− 9.273$^{\beta}$	0.596$^{\beta}$	− 9.572$^{\beta}$	9.572$^{\beta}$	− 0.448$^{\beta}$
FAV(N/OH)-GDYTi	− 14.264	− 13.919	0.344	− 14.091	14.091	− 1.906

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3.11. *Drug release mechanism*

Controlled drug release at the target site is considered one of the most important requirements for an efficient drug delivery system, since it enhances the therapeutic performance of the drug while minimizing undesired side effects on healthy tissues. The relatively high adsorption energies obtained for the investigated complexes suggest strong interactions between Favipiravir and the metal-decorated GDY nanocarrier surface, which is beneficial for efficient drug loading and stabilization. Nevertheless, under physiological conditions, these interactions may become significantly weakened due to external factors such as solvent environment, temperature variations, and particularly acidic pH conditions, thereby facilitating controlled drug release. It is known that the pH of damaged or infected cells is lower than that of normal cells.

To evaluate the pH-responsive release behavior of Favipiravir from the GDY nanocarrier surface, a protonation approach was employed to simulate acidic physiological environments that may arise in intracellular compartments, where the pH is lower than neutral conditions. Such acidic environments are known to influence the protonation state of both drug molecules and nanocarrier surfaces, thereby affecting adsorption strength and facilitating controlled drug release. The effect of protonation on the adsorption was investigated by comparing the interaction distances and bond angles before and after protonation for FAV(N/O)-GDYNi, FAV(N/O)-GDYCu and FAV(N/O)-GDYTi complexes. Protonation was performed on the adsorption site and the nearest interacting atoms between Favipiravir and the decorated GDY surface (92). Following geometry optimization upon protonation, substantial structural changes were observed for three metal-decorated GDY loaded with the drug, Figure 11. Before the protonation, the neutral complexes exhibited shorter interaction distances together with bond angle closer to the ideal linear configuration (180°), indicating strong coordination interactions and effective orbital overlap between Favipiravir and the decorated GDY surfaces. However, after protonation, the adsorption distances become longer and the interaction angles deviated further away from 180°, reflecting distortion of the adsorption geometry and weaking of the coordination interaction.

For the FAV(N/O)-GDYNi complex in particular, the interaction distances increased, and thus weakened, from 1.89 Å and 1.92 Å before protonation to 2.16 Å and 2.01 Å after protonation. In addition, the C-O-Ni bond angle changed from 115.53° to 112.25°, while the C-N-Ni bond angle changed from 131.32° to 119.05° after protonation. Similarly, FAV(N/O)-GDYCu complex exhibits an interaction distance from 1.98 Å and 2.01 Å before protonation to 3.97 Å and 2.27 Å after protonation. In addition, the C-O-Cu

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bond angle changed from 114.66° to 86.62°, whereas the C-N-Cu angle shifted from 129.39° to 95.15° after protonation. For the FAV(N/O)-GDYTi complex, the interaction distances changed from 1.98 Å and 2.26 Å before protonation to 2.24 Å and 2.19 Å after protonation. Furthermore, the C-O-Ti bond angle changed from 124.49 to 93.00°, while the C-N-Ti angle changed from 113.20 to 79.93° after protonation.

The deviation from linearity together with the elongation of the adsorption distances indicate reduced adsorption affinity and lower structural stability of the adsorbed configurations under acidic conditions. This has been further confirmed by obtaining positive binding energy reflecting unfavorable interaction. Consequently, protonation facilitates desorption and release of favipiravir from the GDY surfaces at the target acidic environment.

The obtained findings are consistent with previously reported theoretical investigations on related nanocarrier systems, where protonation under acidic conditions was found to reduce adsorption strength and facilitate drug desorption from carbon-based materials (93). Other reports also indicated that protonation transforms the interaction from strong adsorption to weaker noncovalent attraction accompanied by a noticeable increase in adsorption distance (94, 95). Additionally, several computational studies confirmed that protonation-induced weakening of adsorption interactions is generally associated with lower adsorption energies and enlarged drug–carrier separation distances under acidic conditions (96, 97). Overall, these results confirm that protonation under mildly acidic intracellular conditions effectively weakens the interaction between Favipiravir and the GDY nanocarrier surface, thereby facilitating drug desorption and supporting the feasibility of a pH-responsive drug delivery mechanism within the framework of density functional theory calculations.

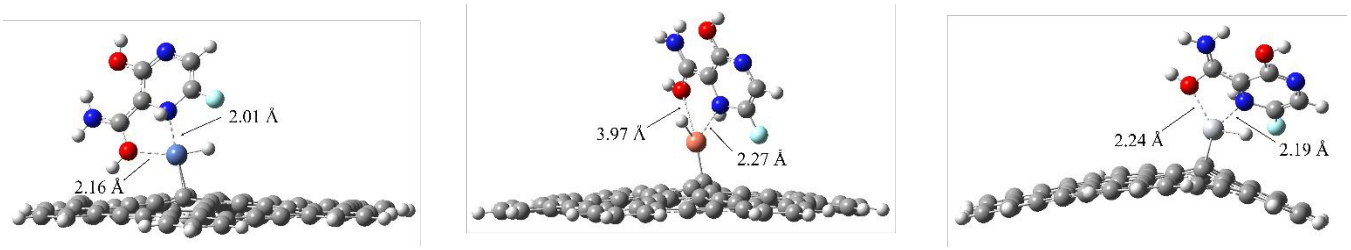

Figure 11. The optimized structures of protonated FAV(N/O)-GDYNi, FAV(N/O)-GDYCu and FAV(N/O)-GDYTi complexes in acidic environment.

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3.12. *Comparison of binding energies (E_{bind}) and binding distances (d) of the FAV drug on pristine and metal decorated GDY nanosheets with other previously reported nanomaterials.*

To evaluate the adsorption performance of favipiravir on pristine and decorated GDY, a comparative assessment performed against reported nanomaterials commonly investigated for drug delivery. This comparison focuses mainly on the calculated adsorption energies and binding distances, which collectively reflect the strength and stability of drug-nanocarrier interactions.

Pristine GDY demonstrates relatively weak adsorption strength, which aligns with the limited interaction capabilities typically reported for undoped carbon-based materials. Upon metal decoration, however, a substantial improvement is observed. All metal-decorated GDY exhibit markedly higher adsorption energies, indicating a more stable and favorable interaction compared with the numerous previously examined materials. This enhancement is consistently evident across the Ni, Cu and Ti modified surfaces, affirming the effectiveness of the decoration approach in reinforcing the adsorption behavior. Among the exanimated variants, Ti decoration provides the most pronounced improvement. These findings demonstrate that metal decoration particularly with Ti substantially enhances the adsorption performance of GDY when assessed against previously investigated nanomaterials.

Table 6. Comparison of binding energies (E_{bind}) and binding distances (d) of the FAV drug on pristine and metal decorated GDY nanosheets with other previously reported nanomaterials, respectively.

System	E_{bind} (eV)	d (Å)	Method	References
FAV				
GDY	**−0.26**	**2.45 Å**	B3LYP/6-31G(d)	**Present work**
GDYTi	**−5.79**	**1.95 Å**	B3LYP/6-31G(d)	**Present work**
GDYNi	**−4.41**	**1.89 Å**	B3LYP/6-31G(d)	**Present work**
GDYCu	**−3.96**	**1.98 Å**	B3LYP/6-31G(d)	**Present work**
C20	−0.17	3.45 Å	M062X/6-31G(d)/BSSE	(98)
C24	−0.085	3.355 Å	B3LYP/6-31G(d,p)	(99)
BNC	−3.469	1.45 Å	ωB97XD/6-31+G*/BSSE	(100)
GNS	−0.296	2.95 Å	PBE/USPP	(49)
BNNS	−0.70	3.43 Å	GGA-PBE/DNP	(101)

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DOI: 10.1039/D6NA00232C

GYNT	−0.85	2.43 Å	B3LYP/6-31G(d)/LANL2DZ/BSSE	(50)
B12N12	−1.120	1.557 Å	B3LYP/6-31G(d,p)	(99)
C23B	−1.434	1.539 Å	B3LYP/6-31G(d,p)	(99)
CB11N12	−1.128	1.555 Å	B3LYP/6-31G(d,p)	(99)
Si-GYNT	−0.74	2.65 Å	6-31G(d)/LANL2DZ/BSSE	(50)
B3O3	−1.30	1.84 Å	ωB97XD/6-31+G(d,p)/BSSE	(102)
BN(Al)NS	−0.88	2.70 Å	GGA-PBE/DNP	(101)
C19Ni	−1.44	1.96 Å	ωB97XD/6-31+G(d,p)/BSSE	(51)
C19Ti	−1.59	2.05 Å	ωB97XD/6-31+G(d,p)/BSSE	(51)
C19Si	−1.88	1.80 Å	BSSE/M062X/6-31G(d)/BSSE	(98)
Be12O12	−1.31	1.67 Å	PBE/DNP	(103)
Mg12O12	−1.52	1.66 Å	PBE/DNP	(103)
Zn12O12	−0.51	1.70 Å	PBE/DNP	(103)
Ni-B12N12	−2.44	-	B3LYP/6–31 G(d,P)	(104)

4. *Conclusion*

In this work, density functional theory (DFT) calculations were employed to systematically investigate the adsorption behavior of favipiravir on pristine and transition metal–decorated graphdiyne (GDY) nanosheets to evaluate their suitability as efficient drug delivery platforms. The structural, energetic, electronic, and thermodynamic characteristics of the resulting drug–carrier complexes were examined in both gas and aqueous environments to assess their stability under idealized and physiologically relevant conditions.

The results demonstrate that pristine GDY exhibits relatively moderate interaction with favipiravir, whereas transition metal decoration significantly enhances the adsorption performance. This enhancement is reflected by larger binding energies, shorter interaction distances, and considerable charge transfer between favipiravir and the modified GDY surfaces. Among the investigated systems, FAV(N/O)–GDYNi, FAV(N/O)–GDYCu, and FAV(N/O)–GDYTi exhibited the strongest adsorption affinities,

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DOI: 10.1039/D6NA00232C

highlighting the important role of metal decoration in tuning the electronic properties and adsorption capability of GDY nanosheets.

Comprehensive electronic analyses, including frontier molecular orbital analysis, density of states, natural bond orbital analysis, reduced density gradient–noncovalent interaction mapping, infrared spectral analysis, global reactivity descriptors, and work function calculations, consistently confirmed the strong interaction between favipiravir and the decorated GDY surfaces. The obtained results indicate that transition metal decoration effectively modulates the electronic structure of GDY, leading to enhanced stability and stronger drug–surface interactions. Thermodynamic calculations further revealed that the adsorption process becomes spontaneous and energetically favorable after metal decoration in both gas and solvent phases, confirming the stability of the investigated systems under realistic conditions.

To further examine the feasibility of controlled drug release, protonated models were investigated to simulate the mildly acidic intracellular environments, where nanocarrier-mediated drug release may occur after cellular internalization. The calculations showed that protonation substantially weakens the interaction between favipiravir and the metal-decorated GDY surface, resulting in reduced binding strength and increased separation distances. These findings suggest that efficient release of favipiravir remains achievable even for strongly adsorbed systems prior to protonation, supporting the pH-responsive behavior of the proposed nanocarriers.

Moreover, a detailed comparison of the optimized geometries and adsorption energies in vacuum and aqueous media revealed only minor changes, indicating that the investigated complexes maintain their structural stability and adsorption characteristics in solvent environments. This observation further supports the tolerance and applicability of the proposed systems under realistic physiological conditions. Overall, this theoretical study demonstrates that transition metal–decorated graphdiyne nanosheets represent promising nanocarriers for favipiravir delivery.

Author contribution

Mohamed M. Aboelnga: Supervision, Project administration, Validation, Computational analysis, Investigation, Conceptualization, Visualization, Formal analysis, Data curation, Software, Resources, Writing-review & editing. **Rana G. Elbayaa:** Writing-original draft, Visualization, Investigation, Formal analysis, Data curation, Investigation, Computational analysis, Methodology. **Elsayed Elbayoumy:**

View Article Online
DOI: 10.1039/D6NA00232C

Supervision, Validation, Conceptualization. **Marco Garavelli:** Formal analysis, Software, Resources, Writing-review & editing. **Mohsen Eltahawy:** Supervision, Validation, Conceptualization, Computational analysis, Investigation, Visualization, Formal analysis, Software, Resources.

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Declaration of Competing Interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

Data availability

Data from this study will be made available from the corresponding author upon reasonable request.

Appendix A. Supplementary material

supplementary data to this article can be found online at

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Data availability

Data from this study will be made available from the corresponding author upon reasonable request.