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Mo₂TiC₂/WSe₂ nanoarchitectures: *in situ* grown nanoflowers for efficient hydrogen electrocatalysis

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The combination of 2D materials provides a powerful strategy to enhance electrocatalytic hydrogen evolution. Efficient hydrogen electrocatalysis is achieved by interfacing conductive Mo₂TiC₂ MXene with catalytic WSe₂ nanoflowers via a one-step hydrothermal route. The hybrid exhibits low overpotential, fast charge transfer, and long-term stability, outperforming pristine components and establishing Mo₂TiC₂/WSe₂ as a promising hydrogen evolution reaction platform.

Introduction

Electrocatalytic water splitting is a key process for sustainable hydrogen production and has attracted significant attention as part of clean energy strategies.^{1,2} The cathodic hydrogen evolution reaction (HER) is particularly important, as platinum remains the benchmark catalyst but is limited by cost and scarcity, motivating the development of earth-abundant alternatives.^{3,4} Two-dimensional materials such as MXenes and transition-metal dichalcogenides (TMDs) have emerged as promising candidates.^{5–10}

Among MXenes, Mo₂TiC₂ has started to attract attention as the combination of molybdenum and titanium provides high conductivity and a more favorable interaction with hydrogen. In addition, the carbide backbone plays an active role by modulating the electronic structure and providing redox-active sites, further facilitating the hydrogen evolution reaction.^{11,12} Although Ti-based MAX phase MXenes have been widely investigated for electrocatalysis, including HER,^{13–15} MXenes derived from Mo-based MAX phases are still relatively

unexplored.^{16–18} The Mo₂TiC₂T_x MXene deriving from the newly synthesized Mo₂TiAlC₂ MAX phase offers a unique layered conductive structure with a large surface area, highly desirable for electrocatalysis. The incorporation of Mo not only enhances hydrogen adsorption properties, bringing the binding energy closer to the optimal range for HER, but also provides redox-active sites (Mo⁶⁺/Mo⁴⁺) that can serve as catalytic centers, adsorption and electron transfer.^{16,19} This synergy with Ti improves electronic conductivity and structural robustness, leading to superior catalytic activity and cycling stability.¹⁹ Moreover, Mo₂TiC₂ has demonstrated promising activity towards electrocatalytic HER, outperforming several other MXenes and even approaching noble metal catalysts,^{20–22} while also serving as an excellent platform for nanoparticle immobilization.^{17,23}

Within the TMD family, WSe₂ has shown promising HER activity, often matching or surpassing MoS₂ when optimized for phase and defect structure.^{24–27} Metallic (1T) WSe₂ exhibits low charge-transfer resistance compared to its semiconducting (2H) phase, and strategies such as defect engineering or doping can further improve proton adsorption and catalytic kinetics.^{26,28,29} These features make WSe₂ a strong candidate for building hybrids with conductive MXenes,¹⁸ where interfacial charge transfer can be further enhanced.

A widely used approach to boost catalytic performance is the *in situ* growth of 2D materials on conductive substrates, which not only provides mechanical stability but also enhances charge transport and reduces interfacial barriers during electrocatalysis. This synergy has been demonstrated in hybrids such as WSe₂/Ti₃C₂Cl₂ in our previous work,³⁰ yet no studies have explored the combination of WSe₂ with Mo₂TiC₂, a newer MXene with potentially stronger interfacial interactions. Mo₂TiC₂'s distinct electronic structure and surface chemistry enable intimate *in situ* growth of WSe₂, enhancing interfacial charge transfer and structural stability—features not observed in the more commonly studied Ti₃C₂-TMD systems. Compared to Ti₃C₂- or V₂C-based MXene/TMD hybrids, the Mo₂TiC₂/WSe₂ system offers unique electronic properties, with Mo⁴⁺/Mo⁶⁺

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redox-active sites³¹ (dominant Mo⁴⁺ with minor Mo⁶⁺ surface contribution),³² facilitating charge transfer and contributing to superior HER performance.

We report the first solvothermal synthesis of WSe₂ nano-flowers directly grown on Mo₂TiC₂ MXene, forming a hybrid with direct interfacial contact. The material shows low overpotential, a small Tafel slope, and improved electrical conductivity, indicating efficient charge transfer and abundant catalytic sites. Stability tests confirm that the Mo₂TiC₂/WSe₂ heterostructure is a durable and effective electrocatalyst for hydrogen evolution.

Experimental

General

All chemicals, reagents, and solvents were purchased from Sigma-Aldrich and used without further purification.

Preparation of Mo₂TiC₂ MXene

The initial powders (Ni, Mo, Ti, Al, C) were sequentially mixed with ethanol to obtain a suitable homogeneous mixture and then dried at 60 °C for 2 hours. Initial composition, of the sample was Ni:Mo:Ti:Al:C = 1:2:1:7:2. Next, 4 g of powder mixtures were pressed (10 kN) uniaxially in the shape of discs (diameter of 25 mm, thickness approx. 3 mm). The samples were placed on the top of the Al₂O₃ powder layer in an alumina crucible and were also covered with an Al₂O₃ powder layer to reduce the evaporation of Al. The powder compacts were sintered to form MAX phase/Ni–Al alloy composite using conventional sintering of initially pressed powders in an Ar atmosphere at 1500 °C and for 4 hours, with a heating/cooling rate of 5 °C per minute, synthesis conditions are described in more detail in the publication.¹⁷

Etching of Al and Ni to form MXene from the sintered and milled sample was carried out in HF solution at 70 °C for 8 days (2 g of powder was added to 40 mL of 50 wt% HF aqueous solution). The etching was carried out in two stages: the first etching lasted 4 days, after which the powder was washed with distilled water to remove AlF₃. The sample was then subjected to a second etching step for an additional 4 days. As a next step, the obtained powder was washed with distilled water (at room temperature), filtered using vacuum filtration with a filter paper and dried at room temperature.

For delamination, TBAOH (20 wt% in water) was used for 1 day. Next, the samples were washed in distilled water to dissolve TBAOH. Samples in the form of powders were filtered through filter paper and dried at room temperature for 1 day.

Preparation of WSe₂

WSe₂ was synthesized by dissolving 1 mmol of tungsten hexacarbonyl and 2 mmol of selenium powder in 30 mL of dimethylformamide (DMF). The mixture was then transferred into a 50 mL Teflon-lined stainless-steel autoclave and heated at 200 °C for 13 hours. Once cooled to room temperature, the suspension was subjected to centrifugation at 10 000 rpm,

followed by washing cycles with DMF (twice), distilled water (three times), and methanol (three times).

Preparation of Mo₂TiC₂/WSe₂

Tungsten hexacarbonyl (175.95 mg) and selenium powder (78.97 mg), along with 3 mg of Mo₂TiC₂ MXene were dissolved in 30 mL DMF and the resulting suspension was transferred into a 50 mL Teflon-lined stainless-steel autoclave reactor and heated at 200 °C for 13 h. After the autoclave was cooled to room temperature, the resulting suspension was centrifuged at 10 000 rpm with DMF (2 times), distilled water (3 times) and methanol (3 times).

Microscopy techniques

The morphology of the synthesized materials was examined by scanning electron microscopy (SEM) using a Tescan Lyra 3 dual microscope with a FEG electron source. Elemental composition and distribution were determined by energy-dispersive X-ray spectroscopy (EDS) equipped with an 80 mm² SDD detector and analyzed with the AZtecEnergy software package (Oxford Instruments). For SEM/EDS analyses, the samples were mounted on carbon-conductive adhesive tape. Transmission electron microscopy (TEM) investigations were carried out with a Jeol EFTEM 2200 FS microscope (Jeol, Japan) operating at 200 kV. Micrographs were captured using a SIS MegaView III digital camera (Soft Imaging Systems) and further processed with AnalySIS v.2.0 software. TEM including high-angle annular dark-field (HAADF) imaging and high-resolution TEM (HRTEM) imaging were obtained in a Zeiss LIBRA 200 MC Cs STEM, operating at 200 kV. For TEM sample preparation, dispersions were made in absolute ethanol, the suspension was then put on carbon lacey TEM grids (200 mesh, PLANO GmbH) by pipette, and dried overnight at room temperature.

XRD

Powder X-ray diffraction measurements were carried out at room temperature using a Bruker D8 Discoverer diffractometer (Bruker, Germany) equipped with a parafocusing Bragg–Brentano geometry and Cu K α radiation ($\lambda = 0.15418$ nm, 40 kV, 40 mA). Patterns were recorded within the 2θ range of 5–70° with a step size of 0.019°. The acquired data were processed using the EVA software package.

XPS

X-Ray photoelectron spectroscopy (XPS) measurements were performed using a Phoibos 100 spectrometer (Specs, Germany) equipped with a monochromatic Al K α 1 source (1486.7 eV). Samples were mounted on Cu conductive tape. High-resolution core-level spectra were acquired with a pass energy of 40 eV and a step size of 0.1 eV. Charge compensation was applied using a flood gun, referencing the C 1s peak to 284.8 eV.

Raman spectroscopy

Raman spectra were collected using an inVia Raman microscope (Renishaw, UK) operated in backscattering geometry and



equipped with a CCD detector. A DPSS laser with 532 nm excitation (50 mW source power, 5 mW applied at the sample) and a 50 \times objective lens was employed. Calibration was performed using a silicon reference with a characteristic peak at 520 cm⁻¹, providing a spectral resolution better than 1 cm⁻¹. For sample preparation, dispersions (1 mg mL⁻¹ in deionized water) were ultrasonicated for 10 min, drop-cast onto silicon substrates, and dried prior to measurement.

Electrochemical measurements for hydrogen evolution reaction

The electrochemical characterization by means of linear sweep voltammetry (LSV) was performed using an Autolab PGSTAT 204 (Metrohm, Switzerland). A standard three-compartment electrochemical cell was used equipped with a rotating disc electrode (RDE) with a glassy carbon disk (geometric surface area: 0.196 cm²) as a working electrode, graphite rod as a counter-electrode, and Hg/HgSO₄ (0.5 M K₂SO₄) as reference electrode. HER LSV measurements were performed at room temperature in Argon-saturated aqueous 0.5 M H₂SO₄ solution. LSV plots were corrected for iR drop by applying a 5–10% iR correction to the measured potentials to account for the ohmic losses, as the correction varied depending on the resistance of the samples. For preparing the catalyst ink, catalytic powder (4.0 mg) was dissolved in a mixture (1 mL) of deionized water, isopropanol, and 5% Nafion (v/v/v = 4 : 1 : 0.02) followed by sonication for 30 min before use. The working electrode was polished with alumina suspension, washed with deionized water, and finally sonicated in double-distilled water before casting 8.5 μ L aliquots of the electrocatalytic ink on the electrode's surface. Finally, electrochemical impedance spectroscopy (EIS) measurements were acquired from 10⁵ to 10⁻¹ Hz with an AC amplitude of 0.01 V. The EIS measurements were conducted at a potential where significant HER current was recorded, corresponding to -2 mA cm⁻².

Results and discussion

The hybrid was prepared *via* a simple solvothermal approach, where WSe₂ formed on the Mo₂TiC₂ MXene, that was used as a platform, through the reaction of tungsten hexacarbonyl with selenium. The Mo₂TiC₂/WSe₂ hybrid was prepared with a WSe₂-rich composition, following our previous optimization on WSe₂/Ti₂C₂Cl₂ hybrids, where higher WSe₂ loading resulted in enhanced HER performance.³⁰ The XRD patterns of Mo₂TiC₂/WSe₂ confirm the successful formation of the hybrid structure. Fig. 1a shows the diffraction patterns of the Mo₂TiC₂/WSe₂ along with the pristine materials, highlighting the diffraction planes corresponding to hexagonal WSe₂ (space group *P*6₃/*mmc*, JCPDS No. 38-1388). The pristine WSe₂ exhibits characteristic reflections at the (002), (100), (103), (105), and (110) planes.³³ In the Mo₂TiC₂/WSe₂ hybrid, the (002) reflection appears less broad compared to the pristine WSe₂, while additional peaks originating from the MXene overlap in the same region (8.57 $^{\circ}$, 11.5 $^{\circ}$). The (100) reflection is downshifted by \sim 1 $^{\circ}$, while the (103) reflection shifts toward higher 2θ by \sim 1 $^{\circ}$, indicating subtle

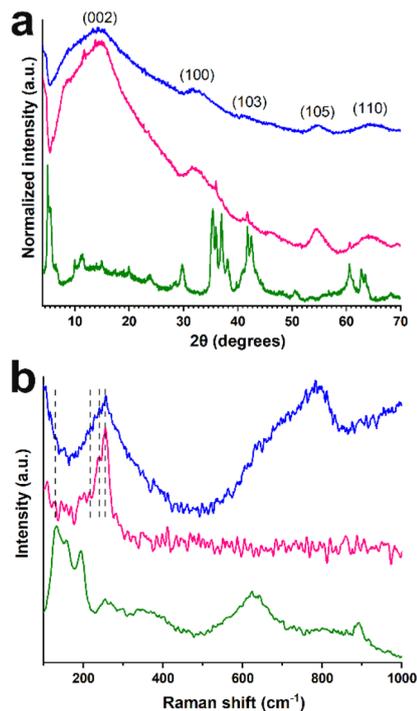


Fig. 1 (a) XRD patterns and (b) Raman spectra for Mo₂TiC₂/WSe₂ (pink), WSe₂ (blue) and Mo₂TiC₂ MXene (green).

lattice distortions associated with interfacial coupling and interlayer strain upon hybrid formation.^{30,34} Additional peaks at 35.9 $^{\circ}$, 37.0 $^{\circ}$, 41.67 $^{\circ}$ and 60.7 $^{\circ}$ originating from the MXene phase, are also detected in the hybrid.³³ XRD patterns of the

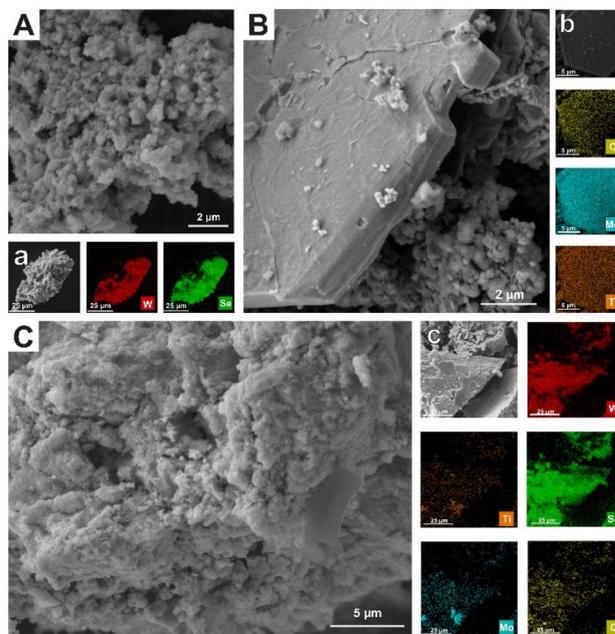


Fig. 2 SEM images of (A) WSe₂ (B) Mo₂TiC₂ MXene and (C) Mo₂TiC₂/WSe₂ hybrid with corresponding EDS elemental maps (a–c) showing the distribution of elements.



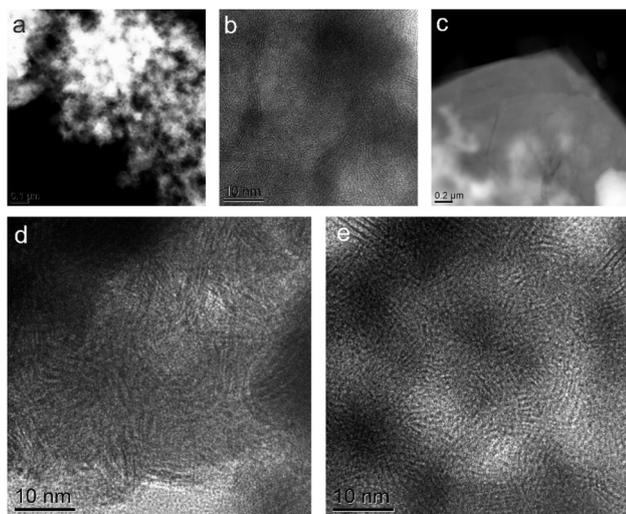


Fig. 3 TEM analysis of WSe₂, Mo₂TiC₂ MXene, and the Mo₂TiC₂/WSe₂ hybrid. (a) HAADF-STEM image and (b) HRTEM image of WSe₂ nano-flowers. (c) HAADF STEM image of a Mo₂TiC₂ MXene flake. (d and e) HRTEM images of the Mo₂TiC₂/WSe₂ hybrid.

synthesized Mo₂TiC₂ MXene are presented in Fig. S1, with the pristine MAX phase included for comparison.¹⁷

The Raman spectra of all materials are shown in Fig. 1b. In the Raman spectrum of pure Mo₂TiC₂T_x, we can ascribe the band at 158 cm⁻¹ to E_g vibrations from both Mo and Ti atoms in the oxygen-terminated Mo₂TiC₂.^{33,35} The E_g vibration at 254 cm⁻¹ can be related to the presence of oxygen. Higher frequencies (360, 420 and 625 cm⁻¹) can be attributed to mostly C vibrations in Mo₂TiC₂.^{33,35} Intact WSe₂ shows bands at 131, 212, and 238 cm⁻¹, corresponding to J₁, J₂, and J₃, which are characteristic of the

metallic 1T octahedral phase of WSe₂.^{30,36} Additionally, the band at 256 cm⁻¹ for WSe₂ results from the overlapping E_{2g}¹ and A_{1g} modes, typical of the 2H phase.^{30,36} In the hybrid material, bands deriving from both individual components are observed: J₁-J₂ bands are seen deriving from WSe₂'s 256 cm⁻¹ band appears downshifted to 254 cm⁻¹ in the hybrid. The latter, along with the observed shift in XRD suggests interfacial coupling between WSe₂ and MXene layers upon hybridization. This region also overlaps with the C vibrations in the MXene.

Scanning Electron Microscopy (SEM) and Transmission Electron Microscopy (TEM) analyses were employed to determine the morphology and elemental composition of the materials (Fig. 2). The SEM image of pristine WSe₂ (Fig. 2A) exhibits its characteristic flower-like aggregates, which is typical of the bottom-up preparation approach. This is also confirmed by HAADF image (Fig. 3a). HRTEM image (Fig. 3b) shows nano size layered crystal features embedded in amorphous structure. In contrast, the MXene displays its typical layered architecture (Fig. 2B and 3c). The corresponding images of the hybrid confirm the coexistence of both components (Fig. 2c). Furthermore, EDS mapping verifies the distribution of the constituent elements throughout the hybrid structure, confirming the successful preparation of the heterostructure (Fig. 2C). HRTEM images (Fig. 3d and e) of the hybrid material show similar nano size layered crystal features embedded in amorphous structure.

Surface composition analysis was performed by X-ray photoelectron spectroscopy (XPS). The XPS survey spectra of the Mo₂TiC₂/WSe₂ hybrid and the reference materials (WSe₂ and Mo₂TiC₂) are shown in Fig. S2. High-resolution (deconvoluted) XPS spectra for WSe₂ are shown in Fig. S3. The high-resolution spectrum of Se 3d consisted of two peaks associated with Se^{II}

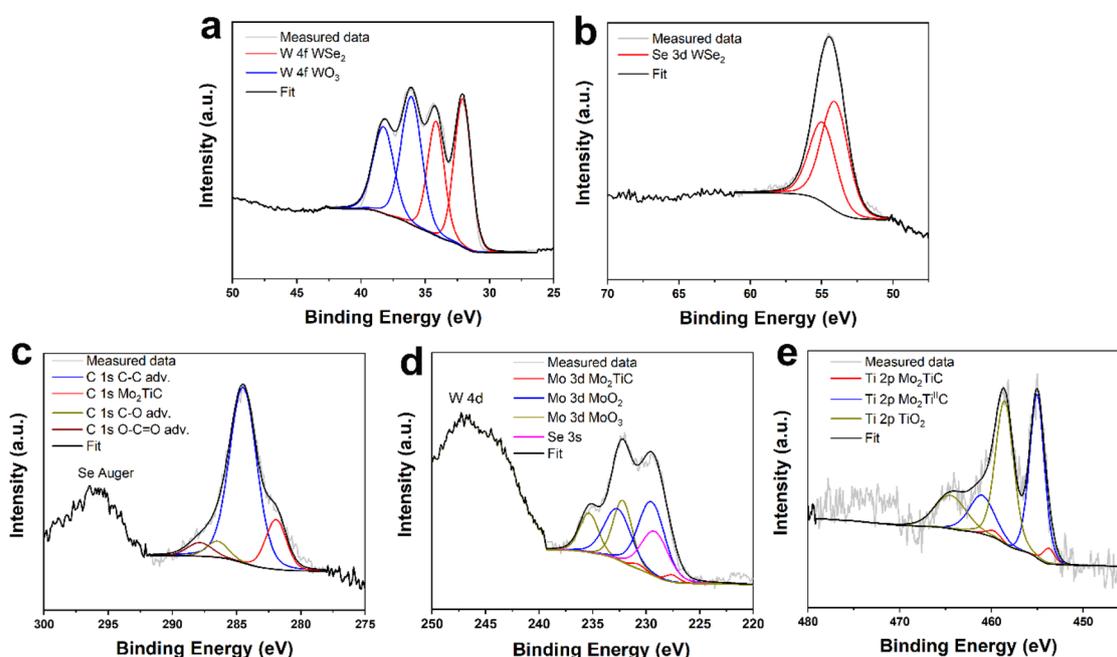


Fig. 4 (a) Deconvoluted X-ray photoelectron spectra of Mo₂TiC₂/WSe₂ showing (a) W 4f and (b) Se 3d, (c) C 1s, (d) Mo 3d and (e) Ti 2p chemical states.



(53.7 and 54.6 eV for Se 3d_{5/2} and 3d_{3/2}, respectively). The W 4f high-resolution spectrum was comprised of two doublets assigned to WSe₂ (31.6 and 33.7 eV for W 4f_{7/2} and W 4f_{5/2}, respectively) and WO₃ (35.5 and 37.7 eV for W 4f_{7/2} and W 4f_{5/2}, respectively). The presence of WO₃ clearly indicates surface oxidation of the material, a phenomenon commonly observed in tungsten-based TMDs and widely reported in the literature. Such surface oxide species have been shown to enhance the electrocatalytic performance of the material.^{25,26,37,38} Mo₂TiC was analyzed in an identical manner, with XPS data shown in Fig. S4. For Mo 3d high-resolution spectrum, an extensive surface oxidation was observed, which is documented by the presence of two doublets originating from MoO₂ (229.6 and 233.0 eV for Mo 3d_{5/2} and Mo 3d_{3/2}, respectively) and MoO₃ (232.5 and 235.7 eV for Mo 3d_{5/2} and Mo 3d_{3/2}, respectively). A doublet associated with the Mo₂TiC MXene phase was also identified in the Mo 3d spectrum (227.5 and 230.8 eV for Mo 3d_{5/2} and Mo 3d_{3/2}, respectively). These results show a high degree of surface oxidation, however, Mo oxides, particularly MoO₂ have been previously associated with good catalytic performance for HER.^{39,40} For the Ti 2p spectrum, three doublets associated with Ti in MXene (454 and 460 eV for Ti 2p_{3/2} and Ti 2p_{1/2}, respectively), slightly oxidized Ti in MXene (455.8 and 462.0 eV for Ti 2p_{3/2} and Ti 2p_{1/2}, respectively) and TiO₂ (459.5 and 465.5 eV for

Ti 2p_{3/2} and Ti 2p_{1/2}, respectively), were identified as previously reported in other reports.⁴¹ Finally, in the C 1s spectrum, a carbide peak at 282.5 eV together with peaks originating from adventitious contamination, were also identified. The XPS spectra of the hybrid material are shown in Fig. 4. A comparison of the high-resolution spectra from this material to the pure parent materials revealed no significant changes in terms of peak positions. Slight changes in the ratios between individual peaks were identified (see Tables S3–S6), but these likely originate from a small degree of variance in the surface composition of the materials. Such results indicate a synergistic effect between both materials, rather than dramatic compositional changes introduced by the synthesis of the hybrid.

Next, the HER performance of Mo₂TiC₂/WSe₂ hybrid electrocatalyst was evaluated, alongside reference materials WSe₂, Mo₂TiC₂ MXene, and 20 wt% Pt/C, using LSV measurements in an Ar-saturated 0.5 M H₂SO₄ aqueous electrolyte (Fig. 5). The Mo₂TiC₂/WSe₂ hybrid shows markedly superior HER activity (Fig. 5a), initiating hydrogen evolution at -0.14 V vs. RHE, which is 120 and 410 mV lower than WSe₂ and Mo₂TiC₂ MXene, respectively. At -10 mA cm⁻², it delivers a low overpotential of 320 mV, outperforming WSe₂ and Mo₂TiC₂ MXene by 160 mV and 380 mV, as the pristine components exhibit substantially higher overpotentials of -0.48 V and -0.70 V vs. RHE.

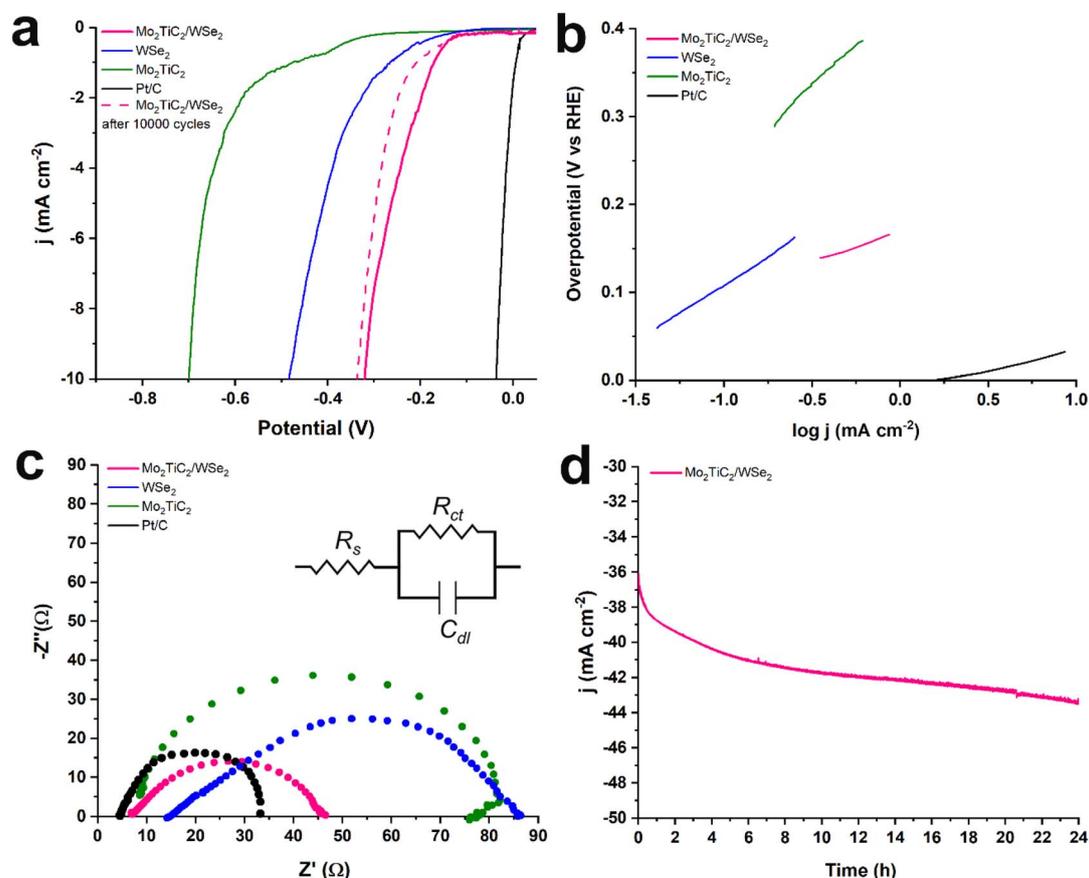


Fig. 5 (a) iR-corrected LSVs for HER obtained at 1600 rpm rotation speed and 5 mV s⁻¹ scan rate before (solid lines) and after 10 000 cycles (dashed lines) in Ar-saturated aqueous 0.5 M H₂SO₄, (b) Tafel slopes, and (c) Nyquist plots for Mo₂TiC₂/WSe₂ (pink), Mo₂TiC₂ MXene (green), WSe₂ (blue), and Pt/C (black). Data were fitted using a Randles equivalent circuit. (d) Chronoamperometric response of Mo₂TiC₂/WSe₂ at -1.56 V (vs. RHE) for 24 h.



The enhanced reaction kinetics of Mo₂TiC₂/WSe₂, corroborated by Tafel and EIS data, can be attributed to the direct contact enabled by the robust deposition of WSe₂ onto Mo₂TiC₂ MXene, facilitating efficient electron transfer within the hybrid structure. Additionally, the abundant flower-like WSe₂ in the hybrid provides a high surface area with more exposed active edge sites for enhanced HER performance.

In order to obtain information about the reaction mechanism, we extracted Tafel slopes from LSV curves (Fig. 5b) and performed electrochemical impedance spectroscopy (EIS) (Fig. 5c). The Mo₂TiC₂/WSe₂ hybrid exhibited the lowest Tafel slope, 74 mV dec⁻¹, showing that the Heyrovsky step is rate-limiting. Protons first adsorb on the surface (Volmer step) and then combine to form H₂. In contrast, pristine WSe₂ and Mo₂TiC₂ MXene had higher slopes, 128 and 191 mV dec⁻¹, pointing to slower kinetics.

EIS at -2 mA cm⁻² further corroborate the above findings by revealing a markedly reduced charge-transfer resistance for the Mo₂TiC₂/WSe₂ hybrid. The hybrid's charge-transfer resistance was 55 Ω, while reference materials displayed higher resistance values of 70 Ω for WSe₂ and 75 Ω for Mo₂TiC₂. This improvement derives from: (i) the direct contact between WSe₂ and Mo₂TiC₂, which helps electron transfer, and (ii) the flower-like WSe₂, which exposes more active sites for HER. The electrochemically active surface area (ECSA) was calculated from the double-layer capacitance (C_{dl}) from CV curves measured in the non-faradaic range (50–500 mV s⁻¹, Fig. S5). In reference capacitance, 40 μF cm⁻² was used to estimate ECSA values, which were 16.5 cm² for Mo₂TiC₂/WSe₂, 1.8 cm² for Mo₂TiC₂, 3.9 cm² for WSe₂. The larger surface area of the hybrid reflects a higher number of accessible active sites. Interestingly, Mo₂TiC₂/WSe₂ exhibited an ECSA value of 10.8 cm² after 10 000 cycles, showing a slight decrease from the original value. Additionally, the specific activity was calculated as $j_{\text{ECSA}} = (j_{\text{geo}} \times A_{\text{geo}}) / \text{ECSA}$ where j_{geo} is the geometric current density and A_{geo} the geometric surface area of the electrode. The Mo₂TiC₂/WSe₂ hybrid shows a lower j_{ECSA} than the pristine materials, indicating that the improvement in HER performance mainly arises from its increased electrochemically active surface area and the high density of exposed active sites rather than intrinsically higher catalytic activity. Table S1 summarizes the electrocatalytic parameters for all tested materials. As summarized in Table S2, the HER performance of the Mo₂TiC₂/WSe₂ hybrid is comparable to other WSe₂- and MXene-based electrocatalysts reported in the literature, supporting its potential as an efficient and stable system.

Furthermore, the stability of the hybrid was evaluated by running 10 000 cycles, as shown in Fig. 5a. After continuous cycling, the hybrid exhibited only a minor potential increase of about 20 mV. To further examine its durability, chronoamperometric measurements were conducted at -1.56 V vs. RHE, corresponding to a current density of approximately -36 mA cm⁻², for 24 h under 1600 rpm (Fig. 5d). The measurement showed an initial activation process, during which the current density progressively increased in magnitude, reaching approximately -43 mA cm⁻² after 24 h of continuous operation. This gradual increase reflects surface activation and demonstrates the excellent long-term

durability of the Mo₂TiC₂/WSe₂ hybrid under prolonged electrolysis conditions. Post-HER SEM imaging reveals that the hybrid retains its overall morphology, with slightly more exposed Mo₂TiC₂ sheets resulting from the partial detachment of loosely bound WSe₂ nanoflowers during gas evolution (Fig. S6). The observed increase in oxygen is attributed mainly to residual Nafion and surface rehydration. No oxide-related structural changes are detected, confirming the stability of both components under HER conditions. Similar behavior is observed in the post-chronoamperometry SEM analysis (Fig. S7), further confirming the structural robustness of the Mo₂TiC₂/WSe₂ hybrid under prolonged electrochemical operation.

Conclusions

In summary, we developed Mo₂TiC₂/WSe₂ nanoarchitectures through a simple solvothermal route, achieving intimate interfacial contact between the Mo₂TiC₂ MXene and WSe₂ nanoflowers. The hybrid exhibited a low overpotential (-0.32 V at -10 mA cm⁻²), a Tafel slope of 74 mV dec⁻¹ consistent with a Volmer–Heyrovsky pathway, and a significantly enlarged electrochemically active surface area. The electrocatalyst exhibited excellent stability, with negligible performance loss after 10 000 cycles and a gradual activation during 24 h chronoamperometry at -36 mA cm⁻², reaching ~-43 mA cm⁻². These results highlight Mo₂TiC₂/WSe₂ as an efficient and durable electrocatalyst for hydrogen evolution.

Conflicts of interest

There are no conflicts to declare.

Data availability

Data analysed during the study are accessible *via* the Zenodo repository <https://zenodo.org/records/17157686>.

Supplementary information (SI): imaging, spectroscopic and electrocatalytic data. See DOI: <https://doi.org/10.1039/d5na01182e>.

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