



Cite this: DOI: 10.1039/d5mh02418h

Received 19th December 2025,
Accepted 1st April 2026

DOI: 10.1039/d5mh02418h

rsc.li/materials-horizons

Mechanistic origin of charge separation and enhanced photocatalytic activity in D- π -A-functionalized UiO-66-NH₂ MOFs

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Donor- π -acceptor (D- π -A) functionalization of MOF linkers can enhance visible-light photocatalytic activity, yet the mechanisms responsible for these effects remain unclear. Here we combine EPR spectroscopy, transient photoluminescence, and first-principles calculations to examine how diazo-coupled anisole, diphenylamine (DPA), and *N,N*-dimethylaniline (NNDMA) groups modify the photo-physics of UiO-66-NH₂. All donor units introduce new occupied states near the valence-band edge, enabling charge separation through dye-to-framework electron transfer. Among them, the anisole-modified material stands out for facilitating efficient inter-system crossing into a triplet charge-transfer configuration that suppresses fast recombination and yields long-lived charge carriers detectable by photo-EPR. Meanwhile, bulkier donors such as DPA and NNDMA – despite their stronger electron-donating character – also tend to introduce defect-associated trap states. These results underscore the interplay between donor-induced electronic-structure changes, triplet pathways, and defect-mediated recombination, offering a mechanistic basis for tuning photocatalytic response in D- π -A-modified MOFs.

Introduction

Metal-organic frameworks (MOFs) have emerged as a versatile class of crystalline porous materials with tunable structures and functionalities, finding applications in areas such as

New concepts

Donor- π -acceptor (D- π -A) functionalization offers a promising route to improve MOF photocatalysts, yet the mechanisms behind these enhancements remain unclear and are often interpreted primarily in terms of optical band-gap narrowing. Here we show that reshaping of excited-state pathways, rather than light harvesting alone, is what governs charge-carrier lifetimes and photocatalytic efficiency in D- π -A-modified MOFs. By combining photo-EPR spectroscopy, transient photoluminescence, and first-principles calculations, we demonstrate that diazo-grafted donors introduce higher-lying charge-transfer (CT) singlet states that act as the entry point for long-range charge separation in UiO-66-NH₂. Crucially, the presence of nearby triplet CT states determines whether charge carriers become spin-protected against fast recombination. Among the series, a moderate donor (anisole) provides an optimal balance, enabling efficient intersystem crossing into a CT triplet state while preserving structural integrity, resulting in long-lived carriers detectable by photo-EPR and superior photocatalytic activity. In contrast, stronger donors promote defect-associated trap states that open additional recombination pathways. These findings establish triplet-enabled charge separation and defect control as key design principles for D- π -A-functionalized MOF photocatalysts.

catalysis, gas storage, and sensing.^{1,2} A notable example is the Zr-based UiO-66 framework,^{3,4} which features Zr₆O₄(OH)₄ clusters linked by terephthalate units and is renowned for exceptional chemical and thermal stability⁵ allowing the inclusion of different linker modifications.⁶ UiO-66 and its derivatives have shown potential in photocatalysis. However, the parent material is a wide-band-gap semiconductor, absorbing mainly in the ultraviolet (UV) range (with a peak around 300 nm)^{7,8} – a significant drawback for solar-driven applications.

One strategy to extend the light absorption of UiO-66 into the visible range is to introduce electron-donating functional groups on its linkers, which introduce new filled (or occupied) states within the band gap and red-shift the absorption edge.^{6,8-10} For example, amino-functionalization reduces the band gap from ~ 4.0 eV in UiO-66 to ~ 2.9 eV in UiO-66-NH₂.⁹

An effective strategy to achieve further enhancement is through the incorporation of organic dyes *via* post-synthetic

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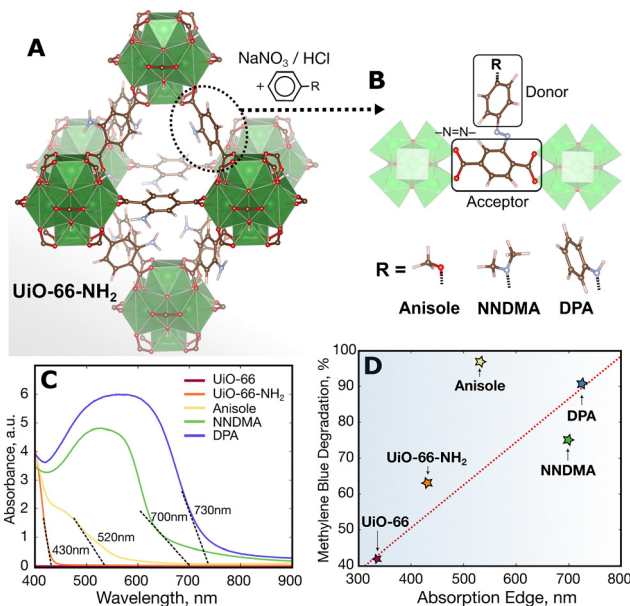


Fig. 1 (A) Schematic structure of UiO-66-NH₂. Color key: carbon—brown; oxygen—red; hydrogen—white; nitrogen—grey; Zr₆O₄(OH)₄ metal nodes—green polyhedra. (B) General chemical structure of D- π -A-like linker functionalization, with an azo ($-\text{N}=\text{N}-$) bond as a π -bridge and different electron-donating groups R (R = anisole, NNDMA, and DPA). (C) Absorption spectra of dye-functionalized MOFs compared to unmodified UiO-66 and UiO-66-NH₂. The spectrum of UiO-66 is limited to the UV range (below 320 nm). (D) Methylene blue degradation versus absorption edge in tested MOFs. For further discussion of data in (C) and (D), see also Ref. 12.

modifications.¹¹ Recent work has shown that post-synthetic covalent modification of UiO-66-NH₂ *via* diazonium coupling can shift the absorption into the visible spectrum.¹² This approach exploits the amino groups of UiO-66-NH₂ as anchoring points, by transforming them into diazo compounds ($-\text{N}\equiv\text{N}^+$) that react with the aromatic moiety of dye molecules forming stable azo ($-\text{N}=\text{N}-$) linkages within the MOF structure¹³ (as illustrated in Fig. 1A and B). While grafting chromophore units onto MOF linkers is recognized as a promising route toward more efficient light harvesting, a detailed understanding of the photophysical consequences is needed to rationalize and optimize the improvements.

In this study, we focus on covalent modification of the UiO-66-NH₂ linkers that follows the design principles of donor- π -acceptor (D- π -A) dyes.¹⁴ Azo dyes with D- π -A architecture contain a $-\text{N}=\text{N}-$ azo bridge connecting an electron-donor aryl group to an electron-acceptor moiety. Such dyes are widely used as photosensitizers to extend the light absorption of photocatalysts (for instance, grafted on TiO₂)¹⁵ and can be tuned by varying their donor, bridge, or acceptor components.¹⁶ A key principle in D- π -A design is to localize the highest occupied molecular orbital (HOMO) on the donor segment and the lowest unoccupied molecular orbital (LUMO) on the acceptor.^{17,18} This promotes intramolecular charge transfer upon photoexcitation, facilitating charge separation: the photoexcited electron shifts from the donor to the acceptor segment and can be then injected into the conduction band of a semiconductor photocatalyst. Implementing this concept in a

MOF, we expect that an appended donor group can act as a visible-light absorber and funnel electrons into the MOF's inorganic nodes, improving light harvesting and photocatalytic activity. By combining experimental and theoretical results, we examine how the modifications influence light absorption, charge separation, and intersystem crossing in the dye-functionalized MOFs and derive insights for the rational design of MOF-based photocatalysts.

Results and discussion

Post-synthetic modifications, optical properties and photocatalytic activity

We investigate UiO-66-NH₂ (Fig. 1A) functionalized with three representative donor aryl groups of increasing electron-donating strength (see Fig. 1B): anisole ($-\text{OC}_6\text{H}_5\text{CH}_3$), *N,N*-dimethylaniline ($-\text{C}_6\text{H}_4\text{N}(\text{CH}_3)_2$), and diphenylamine ($-\text{C}_6\text{H}_4\text{NHC}_6\text{H}_5$). These molecules were covalently attached to the 2-aminoterephthalate linkers of UiO-66-NH₂ *via* diazo coupling (forming $-\text{N}=\text{N}-$ linkages). For brevity, we refer to the resulting MOFs as UiO-66-NH₂-anisole, -NNDMA, and -DPA, respectively. All three organic groups contain strong electron-donating substituents (methoxy, dimethylamino, and diphenylamino), as reflected by their Hammett σ constants in the para position.¹⁹ The terephthalate linker (tethered to Zr nodes) plays the electron acceptor role due to its electron-withdrawing carboxylate groups.

Visually, the MOF powders changed color upon functionalization (from pale yellow for UiO-66-NH₂ to orange or reddish-brown for the dye-functionalized samples), qualitatively indicating successful dye incorporation and enhanced visible light absorption (Fig. S1). The UV-Vis spectra shown in Fig. 1C confirm that all three functionalized MOFs have significantly expanded absorption into the visible region compared to the parent UiO-66-NH₂. The unmodified UiO-66-NH₂ shows a steep absorption edge \sim 430 nm (consistent with its \sim 2.9 eV band gap⁹). Upon grafting the donor groups, the absorption onsets are red-shifted: UiO-66-NH₂-anisole absorbs up to \sim 520 nm, while -NNDMA and -DPA extend further, roughly to 700 nm. This trend aligns with the electron-donating strength of the substituents ($-\text{OCH}_3 < -\text{NHC}_6\text{H}_5 < -\text{N}(\text{CH}_3)_2$): stronger donors push the MOF's frontier orbitals closer together narrowing the gap.

Subsequently, we analyzed the photocatalytic performance of the dye-functionalized MOFs under broadband (UV-visible) illumination (with TiO₂ used as a reference photocatalyst under identical conditions; see Materials and Methods). Fig. 1D illustrates the efficiency in the methylene blue (MB) degradation plotted against the absorption edge obtained from the Tauc plots. All three dye-functionalized MOFs show enhanced photocatalytic performance relative to unmodified UiO-66 and UiO-66-NH₂. Notably, however, the degree of photocatalytic performance enhancement does not strictly follow the extent of visible-light absorption. The UiO-66-NH₂-anisole sample, which had the most modest red-shift in absorption, achieves the highest MB degradation efficiency among the three. This lack of direct correlation suggests that more factors beyond the



shifted absorption edge influence the overall photocatalytic efficiency.

Insights on electronic structure and charge-separation from DFT and PL

To analyze how each functional group modifies the electronic structure of the MOF and to check if this can help rationalize the experimentally observed trends, we carried out DFT calculations. As a benchmark, our DFT calculations for UiO-66-NH₂ yield a band gap of 2.86 eV (Fig. 2A), in excellent agreement with the experimental value.⁹ Next, we computed the electronic gaps of the isolated dye molecules (Fig. 2A). The resulting HOMO–LUMO gaps of anisole, NNDMA, and DPA in the gas phase (calculated for the molecules in vacuum) are all significantly larger (by ~2 eV) than the band gaps of the corresponding functionalized MOFs deduced from the experimental absorption edges (*cf.* Fig. 1C). This indicates that the observed visible absorption in the MOF samples is indeed due to the covalent integration of the dyes into the framework rather than residual free molecules.

Subsequently, we built structural models of dye-modified UiO-66-NH₂ and found that each functionalized linker can be accommodated in a pore of the MOF (*cf.* Fig. S2), although the substituents (especially the larger DPA and NNDMA moieties) do cause some local distortion of the surrounding framework. The phenyl rings of DPA and the dimethylamino group of NNDMA come in close contact with the pore walls causing slight shifts of Zr-node and linker positions (*cf.* Fig. S3).

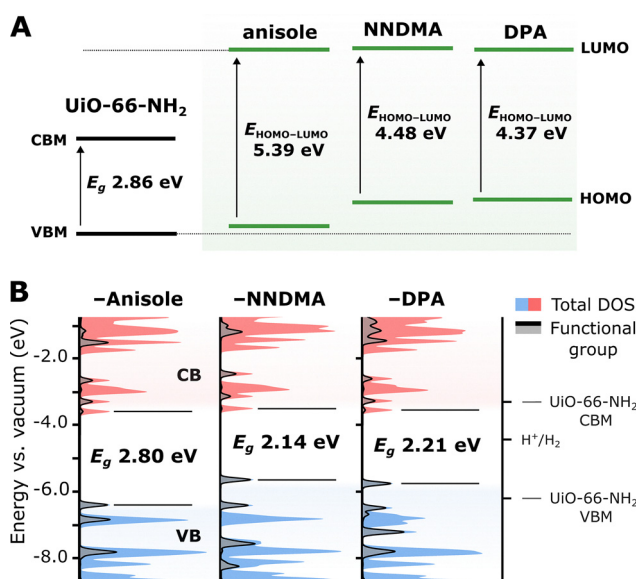


Fig. 2 (A) DFT-calculated band gap of pristine UiO-66-NH₂ compared to the HOMO–LUMO gaps of the free dye molecules in vacuum. (B) Total and projected density of states (DOS) for the functionalized MOFs, showing the emergence of dye-localized states (dark grey area) at the valence band edge and their influence on band gap narrowing. Models assume one dye per pore and partially diazotized linkers bearing OH groups to reflect realistic synthetic conditions. All energies are aligned to the average electrostatic potential in the pore center (vacuum level).

Importantly, our structural models suggest that, due to steric hindrance, a maximum of one dye molecule can be anchored per pore, so the functionalization degree is inherently limited – only one of six linkers can be dye-modified at best. The remaining fraction of diazotized linkers is likely to undergo side reactions. A possible scenario is hydrolysis of the diazonium moiety ($-N\equiv N^+$) to form phenolic $-OH$ on the linker.

This partial functionalization is in good agreement with photoluminescence (PL) spectra of UiO-66-NH₂ and modified structures illustrated in Fig. S4 and S5. While the pristine UiO-66-NH₂ exhibit a single emission peak at 451 nm (with no additional features across the 375–845 nm measurement range), a dual emission character is observed in the dye-modified MOF samples. One PL peak is associated with a fraction of the parent framework, and the second one is attributed to the fraction subjected to the respective linker-functionalization. In UiO-66-NH₂-Anisole, the main peak is red-shifted by about 16 nm and broadened, and the second emission band appears near 600 nm, with an intensity approximately 20% lower than that of the main peak. The UiO-66-NH₂-NNDMA sample displays a 10 nm blue shift of the non-functionalized part along with the appearance of a broad, intense emission at 700 nm. In case of UiO-66-NH₂-DPA, the main peak remains largely unchanged save for slight broadening, while the additional feature appears near 715 nm.

This observation is indicative of partly loaded structures where large parts of the UiO-66-NH₂ structure remain non-modified. The shift and broadening of the UiO-66-NH₂ PL peak can result from strain and structural defects introduced during the functionalization processes. Importantly, even partial functionalization is sufficient for enhanced light-harvesting and photocatalytic performance, as shown in Fig. 1C and D.

To simulate partial functionalization in DFT calculations, the supercell models included five linkers per unit cell with $-OH$ (from diazonium hydrolysis) and one linker is functionalized with a dye. The computed density of states (DOS) for the mixed-linker models illustrated in Fig. 2B reveals that new occupied states appear above the valence band maximum (VBM) of the parent MOF. These states, primarily introduced by the HOMO of the anchored dye molecules, raise the valence band edge upward, hence reducing the gap (by ~0.7 eV for DPA and NNDMA and ~0.1 eV for anisole). The extent of the shift correlates with the donor strength: NNDMA and DPA (with the strongly donating $-N(CH_3)_2$ and diarylamino) introduce higher-lying occupied states than anisole (with the moderate $-OCH_3$ donor).¹⁹ This trend is fully consistent with the bathochromic shifts observed in the UV-Vis spectra (*cf.* Fig. 1C).

Beyond band gap values, DFT provides crucial information on the spatial distribution of the frontier electronic states. Fig. S6 illustrates the charge density associated with the highest occupied orbital (nominal VBM) and the lowest unoccupied orbitals (conduction band minimum, CBM) for the dye-modified MOFs. In all cases, the highest occupied state is mainly localized on the dye-functionalized linker – specifically, over the attached donor molecule and the azo bridge – while the CBM retains the character of unmodified UiO-66-NH₂ and is primarily distributed



over the aryl and carboxylate moieties of the linkers, with partial contribution from the Zr oxide nodes.

This D- π -A type segregation of frontier electronic states is expected to facilitate intraframework charge transfer upon photoexcitation: When the dye-functionalized MOF absorbs a photon, the excited electron transfers from the electron-donating dye moiety to the framework's linker that acts as an acceptor, eventually reaching the catalytic sites (the $Zr_6O_4(OH)_4$ nodes of the MOF). The hole left at the dye segment can be then quenched without causing rapid recombination. To search for signatures of the photoinduced charge transfer we turned to continuous-wave electron paramagnetic resonance (cw EPR) spectroscopy.

Probing photoinduced charge transfer by EPR

Before exploring the effect of photoexcitation, it was important to probe the presence of defects introduced during the functionalization, as these can potentially act as charge trapping sites. The EPR spectra of the MOF samples, collected at room temperature without exposure to light, are shown in Fig. 3A. They reveal the presence of several paramagnetic species.

By simulating the EPR spectra, we identified three primary radical centers. The lowest group of traces in Fig. 3A shows an example of the deconvolution for the UiO-66-NH₂-DPA sample (see Fig. S7 for the simulations of the other functionalized MOFs). The first species is a nitrogen-centered radical, which is present in the parent UiO-66-NH₂ and attributed to an NH[•] group on the linkers.²⁰ It consists of four EPR lines due to hyperfine interaction between the unpaired electron and the nearest ¹⁴N and ¹H nuclear spins. The observation of the NH[•] radical agrees with the observation of two PL peaks in Fig. S4

and indicates the presence of non-modified UiO-66-NH₂ fraction. The other two contributions (marked as Species 1 and 2) that appear only for the dye-modified samples are both described with *g*-factors close to that of the free electron ($g_e = 2.0023$). One contribution is a broad line best fitted assuming a slightly anisotropic *g*-factor ($g_x = 2.003$, $g_y = 2.003$, $g_z = 2.006$), and the other one is a narrow EPR line at $g = 2.004$.

The *g*-factor (or, more specifically, its difference from g_e) is a definitive spectroscopic fingerprint determined by the radical's electronic structure and chemical environment *via* spin-orbit coupling effects. Based on the observed *g* values of Species 1 and 2, we can exclude Zr³⁺ defect centers,^{21,22} which typically exhibit significantly larger *g*-shifts and pronounced anisotropy. The *g*-values close to the free-electron value ($g_e = 2.0023$), together with the absence of strong anisotropy or resolved hyperfine structure, are consistent with delocalized organic radical species. These radicals are likely associated with carbon- or nitrogen-centered species possibly formed from radical chain reactions involved in diazonium coupling^{23,24} or associated with oxygen vacancies on secondary building units,^{25,26} although their exact molecular structure cannot be unambiguously identified.⁷

Importantly, the weights of Species 1 and 2 are relatively low in all three modified MOFs and comparable to that of the NH[•] radical. The latter is an intrinsic defect present in UiO-66-NH₂ in low concentrations (the order of 10⁻⁶ spins per linker, as identified in our previous work²⁰). This indicates that the diazo-coupling reaction did not generate a high concentration of defects, and the framework's integrity remains intact. This is further confirmed by powder X-ray diffraction (pXRD) data, showing that the diffraction patterns of the functionalized samples remain consistent with that of the reference

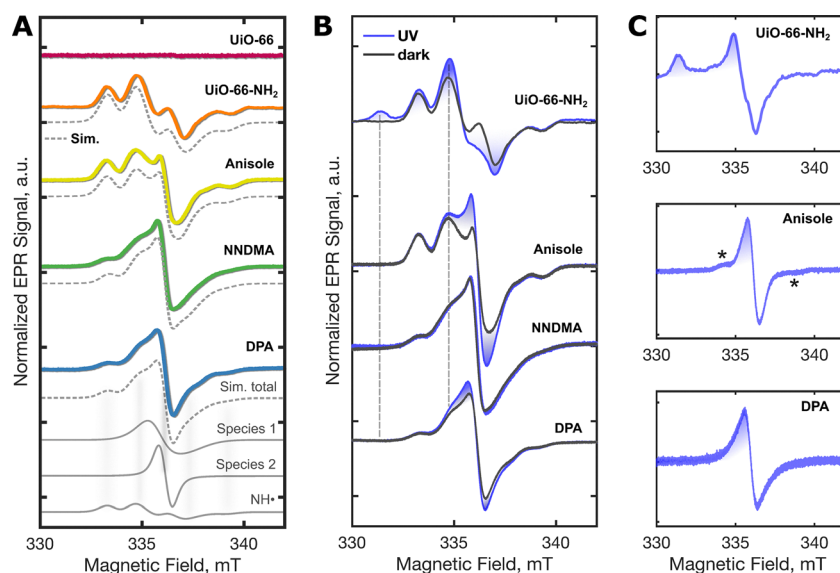


Fig. 3 (A) EPR spectra (solid lines) measured at room temperature in the dark together with spectral simulations (dotted). The parent MOF UiO-66 has no EPR signal, while we detect a strong NH[•] related signal for all modified MOFs. For DPA the individual signal contributions are shown in the lowest group of spectra. All spectra can be simulated as a superposition of the NH[•] group and two unspecified organic radicals, species 1 and 2. (B) Upon broadband UV-irradiation (blue traces) compared with the dark spectra (gray traces) additional signal intensity emerges. Field positions where the main photoinduced changes in the UiO-66-NH₂ EPR spectrum occur, are marked with vertical dashed lines. (C) Photoinduced signals obtained as a difference between the illuminated and dark EPR spectra.



UiO-66-NH₂ framework and no significant broadening of the peaks is observed (Fig. S8). Among the three functionalized samples, UiO-66-NH₂-anisole exhibits the smallest concentration of Species 1 and 2 (*cf.* Table S1 and Fig. S7). This may suggest that the smaller anisole group bonds more cleanly with less perturbation to the structure, whereas the bulkier DPA and NNDMA introduce slightly more disorder (*e.g.* occasional incomplete grafting, side-reactions leading to radicals, and the generation of missing-linker defects^{27–33}). Overall, the EPR characterization confirms that while a few radical species are present after functionalization, their concentrations are reasonably low.

Now to elucidate the photoinduced charge transfer mechanisms in the dye-modified MOFs, we investigated the formation of electron spin centers under *in situ* illumination. Photoinduced spin centers are a telltale sign of electron–hole separation associated with the MOF's photocatalytic activity.⁷ Fig. 3B shows EPR spectra in the dark together with those taken under constant UV-rich illumination (280–450 nm). Changes in the EPR spectra can be observed at several spectral positions. They disappear after UV illumination is switched off, excluding their connection to photoionization damage of the material. To better highlight these changes, Fig. 1C shows the spectra with the dark contributions subtracted.

In the parent UiO-66-NH₂ MOF, the photoinduced EPR signal (Fig. 1C, top panel) is well characterized and commonly attributed to a superoxide (O₂^{•−}) radical formed due to electron transfer to an O₂ molecule adsorbed at an uncoordinated Zr site.^{7,8} It exhibits a nearly axially symmetric *g*-tensor with the principal values $g_x = 2.002$, $g_y = 2.009$, and $g_z = 2.033$. Similar to the parent framework, the anisole-functionalized sample shows appreciable photo-EPR response, suggesting efficient electron–hole separation. At the same time, its photoinduced EPR spectrum is substantially different. It consists of a dominant line at $g \approx 2.004$, accompanied by broad, asymmetric satellite

features near $g \approx 2.014$ and $g \approx 1.990$ (marked with asterisks in Fig. 3C). Most likely, chemical conditions involved in diazo functionalization suppress the formation of node-bound superoxide in favor of other spin centers. Also remarkable is that the photoinduced EPR signal in the DPA sample is different consisting of only one broad line.

One possible explanation for the differences between the anisole- and DPA-modified samples is that their photo-EPR spectra at least partially originate from spin centers localized on the dye-modified linkers, which represent the primary structural difference between the two frameworks. Since the functional groups act as donors, a plausible scenario involves the formation of a cation radical associated with a photogenerated hole following charge separation. To explore this possibility, we modelled such cation radicals using DFT, calculated their EPR parameters, and simulated the corresponding EPR spectra (Fig. S9). The results show that the spatial distribution of spin density can qualitatively account for the observed spectral differences. In the anisole case, the spin density substantially localizes on the diazo π -bridge, producing sizeable hyperfine coupling with the ¹⁴N nuclei and giving rise to the splitting consistent with the satellite features observed experimentally. In contrast, for the DPA-modified linker the spin density is more delocalized over the aromatic donor unit, reducing the hyperfine interactions and yielding a simulated spectrum that remains within the broad experimental linewidth.

Another notable observation is that the DPA-modified MOF exhibits a considerably weaker photoinduced EPR signal than the anisole sample, while NNDMA showed no discernible photo-EPR response (Fig. 1C, middle and bottom panels). This was an unexpected result, since our electronic structure calculations suggest that all three dye-modified MOFs should exhibit efficient charge transfer from the electron-donating functional group to the framework. To seek for a possible explanation, we

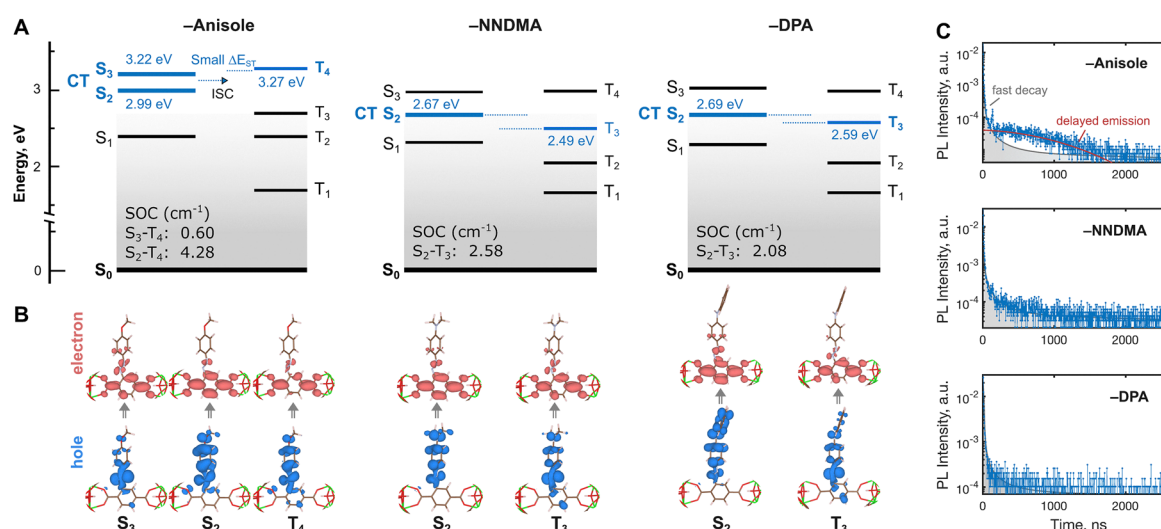


Fig. 4 (A) TDDFT-derived excited-state structure of UiO-66-NH₂ modified with anisole, NNDMA, and DPA donors. The charge-transfer (CT) singlet and triplet states are highlighted in blue, and the calculated spin–orbit coupling (SOC) matrix elements between them are shown in each panel. (B) Electron (red) and hole (blue) components of the difference orbitals associated with the identified CT states. (C) Time-resolved PL traces showing a fast decay of singlet excitation. For the anisole-modified MOF this is followed by delayed emission from triplet states.



carried out time-dependent DFT (TDDFT) calculations accompanied by transient photoluminescence (trPL).

At the first glance, the TDDFT results look similar across all three dye-modified systems (Fig. 4A and B). The lowest singlet excited states (S_1) are of localized exciton (LE) character associated with the $-N=N-$ bridge, while the actual charge-transfer (CT) states that generate charge-separated spins sit between 2.6 and 3.2 eV (≈ 480 – 390 nm), which is still accessible by the excitation wavelengths used in our EPR experiments. At the same time, the anisole-functionalized system exhibits a CT triplet state T_4 that is nearly resonant with one of the CT singlets, S_3 (with the singlet–triplet energy gap $\Delta E_{ST} = 0.05$ eV). Along with the sizeable spin–orbit coupling (SOC) connecting these states, this suggests an efficient intersystem crossing (ISC) pathway. Once the system reaches a triplet CT state, the charge-separated donor–acceptor pair ($D^{\bullet+}/A^{\bullet-}$) becomes spin-protected: recombination is spin-forbidden, allowing the carriers to persist longer, find trapping sites, and yield a detectable photo-EPR signal.

In the NNDMA and DPA systems, the calculations suggest somewhat larger singlet–triplet gaps (ΔE_{ST}), which might be enough to suppress ISC and facilitate faster recombination of the charge carriers. The lifetimes of $D^{\bullet+}/A^{\bullet-}$ pairs in these MOFs can still be sufficient for enhanced photocatalytic activity compared to the parent MOF (*cf.* Fig. 1D) but too short to yield a steady EPR signal at room temperature.

These conclusions agree with the trPL measurements shown in Fig. 4C (for detailed trPL results, see also Fig. S10). The main UiO-66-NH₂ emission band exhibits a similar decay behaviour for all samples, confirming comparable decay dynamics of the host structure. However, the additional linker-related emission bands show distinct characteristics. The anisole-modified MOF sample displays a fast decay component ($\tau_1 \approx 0.4$ ns) followed by a long-lived delayed emission extending up to 1500 ns, indicating phosphorescence from triplet excited states. This dual-component decay – fluorescence and delayed phosphorescence – is also observed for the parent UiO-66-NH₂ MOF (*cf.* Fig. S10). In contrast, long-lived triplet emission is not observed for UiO-66-NH₂-NNDMA and UiO-66-NH₂-DPA samples, whose linker-related emissions are dominated by fast singlet-based radiative decay pathways.

Together the mechanistic insights from EPR, TDDFT, and trPL may explain why anisole functionalization leads to superior photocatalytic performance compared to more electron-donating DPA and NNDMA.

Conclusion

Our results have shown that grafting D– π –A dye units onto the linkers of UiO-66-NH₂ *via* diazo coupling offers a viable route to extend light absorption and enhance photocatalytic activity. Functionalization with anisole, DPA, and NNDMA systematically modifies the MOF's electronic structure by introducing new occupied states near the valence band edge. DFT calculations confirm that these states enable spatial charge separation *via* electron transfer from the dye to the framework backbone,

while preserving the integrity of the host. EPR spectroscopy confirmed that only modest content of defect species is present in the dye-modified samples (indicating the framework remains largely intact).

By combining EPR spectroscopy with trPL measurements and TDDFT calculations, we find that the anisole functionalization strikes an optimal balance. It facilitates efficient charge transfer along with ISC into the triplet manifold that protects the photoinduced charge carriers from fast recombination. Moreover, it appears to introduce fewer defect-related paramagnetic centers (potential trap states) than the bulkier donor groups, thus preserving high charge-transfer efficiency.

From a practical standpoint, the insights gained here provide guidelines for future design of D– π –A-functionalized MOF photocatalysts. We highlight that involving characterization techniques like EPR is imperative for diagnosing trap states, charge-transfer efficiency, and the nature of reactive intermediates. More generally, our work contributes to the growing toolbox for engineering MOF-based materials for solar energy conversion and environmental remediation, connecting concepts from molecular dye chemistry with the robust scaffolds of reticular frameworks.

Methods

Synthesis of dye-functionalized MOFs

UiO-66-NH₂ was synthesized according to Ref.³⁴ and subsequently post-synthetically modified *via* a diazonium coupling reaction to attach the aryl donor groups, following a similar approach to Otal *et al.*¹²

Diffuse reflectance spectroscopy

Reflectance spectra were measured using a UV-visible scanning spectrophotometer (SHIMADZU UV-2101PC) attached to an integrating sphere. Sample powders were mounted into the integrating sphere using a quartz slide. Reflectance was transformed to absorbance using the Kubelka–Munk model.

Photocatalytic characterization

Photocatalytic activity was evaluated using the degradation of methylene blue (MB) as a model reaction. A stock solution of MB (2 mM) was prepared by dissolving 0.0639 g of dye in 100 mL of distilled water. This solution was diluted 1:10 to yield a working concentration of 0.2 mM. For each test, 15 mg of MOF sample was dispersed in 10 mL of the MB solution in a Pyrex test tube equipped with a magnetic stir bar. Titanium dioxide (TiO₂, anatase phase, Alfa Aesar, surface area 45 m² g^{−1}) was used as a reference photocatalyst. Prior to illumination, all suspensions were stirred in the dark for 1 hour to assess dye adsorption. After this equilibration step, 3 mL aliquots were withdrawn, filtered through 20 μ m nylon membranes, and the absorbance was measured at 664 nm to determine the initial MB concentration. The remaining suspensions were then irradiated for 1.5 hours under continuous stirring using a 300 W xenon arc lamp as the light source. After irradiation, aliquots



were again collected, filtered, and analyzed spectrophotometrically at 664 nm.

EPR Spectroscopy

CW EPR was done on a Magnetech MS5000 outfitted with an Oxford ESR 900 He flow cryostat. The deviation of the set temperature was less than 0.2 K. Two types of light sources were used, a broadband Dymax BlueWave 50 UV lamp (280 nm to 450 nm). The power of the UV was measured considering the distance between the sources and the irradiated sample. Thus, the broadband UV lamp has a power of 5.1 mW, respectively. A microwave power of 1 mW was chosen for optimal signal-to-noise ratio of the main EPR spectra without saturation effects. However, experiments on saturation behavior have also been conducted (Fig. S9). EPR spectral simulations were performed with the EasySpin software package³⁵ for MATLAB.

Computational details

The atomic structures and electronic properties of both unmodified and dye-functionalized UiO-66-NH₂ were investigated using density functional theory (DFT) under periodic boundary conditions. All DFT calculations were performed with the Quantum ESPRESSO package v7.2.^{36,37} The generalized gradient approximation of the PBE exchange–correlation functional³⁸ was used for geometry optimizations, with a plane-wave kinetic energy cutoff of 700 eV, Brillouin zone sampling at the Γ -point, along with PAW pseudopotentials taken from PSLibrary v1.0.³⁹ Starting from the experimentally reported UiO-66-NH₂ crystal structure, we constructed supercell models incorporating one functionalized linker (with anisole, NNDMA, or DPA) per Zr₆ node (*i.e.*, one dye per pore) along with five remaining linkers in a partially diazotized state (modeled as –OH substituents on what were formerly –NH₂ groups) to simulate incomplete reactions and potential side-products. The atomic positions were fully relaxed until forces converged below 0.01 eV Å^{–1}, while the lattice constants were kept fixed. Electronic structure calculations were then carried out using the screened hybrid HSE06 functional on the optimized structures to obtain accurate band gap estimates.^{40,41} To align the energy levels and compare band edge positions, we evaluated the electrostatic potential in the center of the MOF pores and took its average as the vacuum level reference, following the procedure of Butler *et al.*⁴² This was facilitated by the MacroDensity tool.⁴³ TDDFT calculations were performed with the ORCA program package⁴⁴ using cluster models cut from the DFT optimized supercells. The clusters consisted of one linker and two Zr-oxo nodes; the remaining (unsaturated) coordination sites of each Zr-oxo node were terminated by HCOO[–] anions. The TDDFT calculations employed the TPSSh exchange–correlation functional⁴⁵ and the def2-TZVP basis set⁴⁶ combined with the RIJCOSX approximation and the corresponding auxiliary basis set.⁴⁷

Photoluminescence and time-resolved photoluminescence measurements

For photoluminescence (PL) and time-resolved photoluminescence (trPL) measurements, powder samples were placed inside

the 0.5 mm thick metal frame sealed between two glass slides (transparent in the 320–2500 nm range) from both sides (Fig. S1). The PL and trPL measurements were performed using an Edinburgh Instruments FLS980 spectrometer. The samples were optically excited with a pulsed 375 nm laser source.

Powder X-ray diffraction (pXRD)

Diffraction patterns were recorded in Bragg–Brentano geometry using PANalytical Empyrean (Cu K α radiation, $\lambda = 1.54056 \text{ \AA}$) and Philips PW3710 (Co K α radiation, $\lambda = 1.78901 \text{ \AA}$) diffractometers.

Author contributions

A.K. performed the EPR and photo-EPR measurements and simulations. V.V. carried out the photoluminescence and time-resolved photoluminescence measurements. A.K., V.V., and A.S. analyzed and interpreted the spectroscopic data. T.B. and W.G.S. performed the first-principles calculations and theoretical analysis. E.O. and K.T. conducted the chemical synthesis, photocatalytic experiments, and UV–Vis characterization. A.K. and T.B. conceptualized and coordinated the project. All authors contributed to the discussion of the results and to writing the manuscript.

Conflicts of interest

There are no conflicts to declare.

Data availability

All data supporting the findings of this study are available within the article and its supplementary information (SI). Supplementary information is available. See DOI: <https://doi.org/10.1039/d5mh02418h>.

Additional data related to this paper may be requested from the corresponding authors.

Acknowledgements

A.K., A.S. acknowledge financial support from the Würzburg-Dresden Cluster of Excellence on Complexity and Topology in Quantum Matter ct.qmat (EXC 2147, DFG project ID 390858490). T.B. and W.G.S. acknowledge the Paderborn Center for Parallel Computing (PC2) for the provided computational resources. The authors thank Manuela L. Kim for assistance with the photocatalytic characterization and Mauricio Calvo for assistance with diffuse reflectance measurements.

References

- H. Furukawa, K. E. Cordova, M. O’Keeffe and O. M. Yaghi, *Science*, 2013, **341**, 1230444.
- H.-C. Zhou, J. R. Long and O. M. Yaghi, *Chem. Rev.*, 2012, **112**, 673–1268.



- 3 J. H. Cavka, S. Jakobsen, U. Olsbye, N. Guillou and C. Lamberti, *et al.*, *J. Am. Chem. Soc.*, 2008, **130**, 13850–13851.
- 4 C. Gomes Silva, I. Luz, F. X. Llabrés i Xamena, A. Corma and H. García, *Chem. Eur. J.*, 2010, **16**, 11133–11138.
- 5 K. Leus, T. Bogaerts, J. De Decker, H. Depauw and K. Hendrickx, *et al.*, *Microporous Mesoporous Mater.*, 2016, **226**, 110–116.
- 6 M. Kandiah, M. H. Nilsen, S. Usseglio, S. Jakobsen and U. Olsbye, *et al.*, *Chem. Mater.*, 2010, **22**, 6632–6640.
- 7 A. Kultaeva, T. Biktagirov, A. Sperlich, P. Dörflinger, M. E. Calvo, E. Otal and V. Dyakonov, *Adv. Funct. Mater.*, 2024, **35**, 2413297.
- 8 J. Long, S. Wang, Z. Ding, S. Wang, Y. Zhou, L. Huang and X. Wang, *Chem. Commun.*, 2012, **48**, 11656–11658.
- 9 E. Flage-Larsen, A. Røyset, J. H. Cavka and K. Thorshaug, *J. Phys. Chem. C*, 2013, **117**, 20610–20616.
- 10 K. Hendrickx, D. E. P. Vanpoucke, K. Leus, K. Lejaeghere and A. Van Yperen-De Deyne, *et al.*, *Inorg. Chem.*, 2015, **54**, 10701–10710.
- 11 Z. Wang and S. M. Cohen, *Chem. Soc. Rev.*, 2009, **38**, 1315–1329.
- 12 E. H. Otal, M. L. Kim, M. E. Calvo, L. Karvonen, I. O. Fabregas, C. A. Sierra and J. P. Hinestroza, *Chem. Commun.*, 2016, **52**, 6665–6668.
- 13 J. Gascon, A. Corma, F. Kapteijn and F. X. Llabrés i Xamena, *ACS Catal.*, 2014, **4**, 361–378.
- 14 J. N. Clifford, E. Martínez-Ferrero, A. Viterisi and E. Palomares, *Chem. Soc. Rev.*, 2011, **40**, 1635–1646.
- 15 Y. Ma, X. Wang, Y. Jia, X. Chen, H. Han and C. Li, *Chem. Rev.*, 2014, **114**, 9987–10043.
- 16 B. G. Kim, K. Chung and J. Kim, *Chem. – Eur. J.*, 2013, **19**, 5220–5230.
- 17 J. L. Brédas, D. Beljonne, V. Coropceanu and J. Cornil, *Chem. Rev.*, 2004, **104**, 4971–5004.
- 18 D. P. Hagberg, T. Marinado, K. M. Karlsson, K. Nonomura and P. Qin, *et al.*, *J. Org. Chem.*, 2007, **72**, 9550–9556.
- 19 C. Hansch, A. Leo and R. W. Taft, *Chem. Rev.*, 1991, **91**, 165–195.
- 20 T. Biktagirov, E. Otal, L. Trost, P. Dörflinger, K. Teshima, O. Trukhina, D. Klose, W. G. Schmidt and A. Kultaeva, *J. Phys. Chem. C*, 2025, **129**, 22210–22218.
- 21 C. Gionco, M. C. Paganini, E. Giamello, R. Burgess, C. Di Valentin and G. Pacchioni, *Chem. Mater.*, 2013, **25**, 2243–2253.
- 22 M. J. Torralvo, M. A. Alario and J. Soria, *J. Catal.*, 1984, **86**, 473–476.
- 23 S. H. Lohmann, K. J. Trerayapiwat, J. Niklas, O. G. Poluektov, S. Sharifzadeh and X. Ma, *ACS Nano*, 2020, **14**, 17675–17682.
- 24 G. Schmidt, S. Gallon, S. Esnouf, J. P. Bourgoïn and P. Chenevier, *Chem. – Eur. J.*, 2009, **15**, 2101–2110.
- 25 R. Xu, Q. Ji, P. Zhao, M. Jian, C. Xiang, C. Hu, G. Zhang, C. Tang, R. Liu, X. Zhang and J. Qu, *J. Mater. Chem. A*, 2020, **8**, 7870–7879.
- 26 S. Tu, A. Liu, H. Zhang, L. Sun, M. Luo, S. Huang, T. Huang and H. Peng, *Chin. Chem. Lett.*, 2024, **35**, 109761.
- 27 W. Xiang, Y. Zhang, Y. Chen, C. J. Liu and X. Tu, *J. Mater. Chem. A*, 2020, **8**, 21526–21546.
- 28 Y. Feng, Q. Chen, M. Jiang and J. Yao, *Ind. Eng. Chem. Res.*, 2019, **58**, 17646–17659.
- 29 B. Bueken, N. Van Velthoven, A. Krajnc, S. Smolders and F. Taulelle, *et al.*, *Chem. Mater.*, 2017, **29**, 10478–10486.
- 30 H. Wu, Y. S. Chua, V. Krungleviciute, M. Tyagi, P. Chen, T. Yildirim and W. Zhou, *J. Am. Chem. Soc.*, 2013, **135**, 10525–10532.
- 31 A. De Vos, K. Hendrickx, P. Van Der Voort, V. Van Speybroeck and K. Lejaeghere, *Chem. Mater.*, 2017, **29**, 3006–3019.
- 32 J. K. Bristow, K. L. Svane, D. Tiana, J. M. Skelton, J. D. Gale and A. Walsh, *J. Phys. Chem. C*, 2016, **120**, 9276–9281.
- 33 M. Vandichel, J. Hajek, F. Vermoortele, M. Waroquier, D. E. De Vos and V. Van Speybroeck, *CrystEngComm*, 2015, **17**, 395–406.
- 34 A. Schaate, P. Roy, A. Godt, J. Lippke, F. Waltz, M. Wiebcke and P. Behrens, *Chem. – Eur. J.*, 2011, **17**, 6643–6651.
- 35 S. Stoll and A. Schweiger, *J. Magn. Reson.*, 2006, **178**, 42–55.
- 36 P. Giannozzi, S. Baroni, N. Bonini, M. Calandra, R. Car and C. Cavazzoni, *et al.*, *J. Phys.: Condens. Matter*, 2009, **21**, 395502.
- 37 P. Giannozzi, O. Andreussi, T. Brumme, O. Bunau and M. B. Nardelli, *et al.*, *J. Phys.: Condens. Matter*, 2017, **29**, 465901.
- 38 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865.
- 39 A. Dal Corso, *Comput. Mater. Sci.*, 2014, **95**, 337–350.
- 40 J. Heyd, G. E. Scuseria and M. Ernzerhof, *J. Chem. Phys.*, 2003, **118**, 8207–8215.
- 41 J. Heyd, G. E. Scuseria and M. Ernzerhof, *J. Chem. Phys.*, 2006, **124**, 219906.
- 42 K. T. Butler, C. H. Hendon and A. Walsh, *J. Am. Chem. Soc.*, 2014, **136**, 2703–2706.
- 43 MacroDensity (software), <https://github.com/WMD-group/MacroDensity> (accessed 2025).
- 44 F. Neese, F. Wennmohs, U. Becker and C. Riplinger, *J. Chem. Phys.*, 2020, **152**, 224108.
- 45 V. N. Staroverov, G. E. Scuseria, J. Tao and J. P. Perdew, *J. Chem. Phys.*, 2003, **119**, 12129–12137.
- 46 F. Weigend, *Phys. Chem. Chem. Phys.*, 2006, **8**, 1057–1065.
- 47 F. Neese, F. Wennmohs and A. Hansen, *J. Chem. Phys.*, 2009, **130**, 114108.

