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## The catalytic conversion of fructose to difructose anhydride

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Difructose anhydride (DFA) is a highly valuable compound, traditionally obtained from inulin or sucrose through enzymatic routes. This work reports a novel eco-efficient process for DFA production from abundantly available fructose in a biomass-derived solvent over a commercial Brønsted acidic beta zeolite. A systematic evaluation of the reaction conditions led to the observation that gamma-valerolactone (GVL) is the most selective solvent giving a DFA yield of 75% under mild reaction conditions. The addition of water as a co-solvent (to improve fructose solubility) suppresses the catalytic activity. Reagent and solvent partitioning was investigated using ssNMR, utilizing the residual dipolar couplings of adsorbed species in the zeolite pores in contrast to the bulk environment. By utilizing a cross-polarization pulse sequence to observe these adsorbed species and a direct polarization to observe all species in the sample, we obtained the ratio of the adsorbed species to overall species in each sample at varying water content in the reaction mixture. Using this approach, we observed that at the  $\geq 10$  volume% water content mark, fructose is no longer able to enter the zeolite pores, coinciding with reaction conditions where DFA is no longer produced. The results of this study illustrate the importance of substrate and solvent partitioning on liquid phase reactions over microporous catalysts.

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### Green foundation

1. High-fructose-corn syrup is a major biomass-based product. This sweetener is however known to have adverse health effects. In this work, we report on an efficient catalytic process to convert fructose into difructose anhydride (DFA) over Brønsted acidic beta zeolite. DFA is also a sweetener itself but has many health benefits. Currently, DFA is only available *via* expensive enzymatic routes from inulin or sucrose making it too expensive for large scale production. In other words, we provide a much more viable pathway with a lower footprint from a readily available source.
2. We obtained DFA yields of 75% using mild reaction conditions using biomass-derived  $\gamma$ -valerolactone as solvent and a stable green catalyst. The main by-product is water.
3. The reaction demonstrates fundamentally interesting solvent effects in that the fructose uptake is highly dependent on the solvent composition. The method that we demonstrate in this work to investigate the partitioning of compounds (inside/outside the porous catalyst) should help further develop this or other liquid phase reactions over microporous catalysts.

## Introduction

Monosaccharides such as glucose and fructose can be readily derived from the cellulosic fraction of biomass and converted

into value-added chemicals. Fructose is a widely available raw material since it is easily obtained from sugar cane,<sup>1</sup> corn processing,<sup>2</sup> inulin<sup>3</sup> and levan depolymerization,<sup>4</sup> as well as from glucose isomerization.<sup>5–11</sup> Owing to its oxyfunctionalized

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nature, fructose serves not only as a key feedstock for the production of potential platform chemicals such as 5-hydroxymethylfurfural (HMF), but also as a promising precursor for high-value compounds used in pharmaceuticals, personal care products, and food additives.<sup>12</sup> Currently, HMF production is the most studied acid-catalyzed reaction for upgrading fructose that employs Brønsted acid catalysts.<sup>13–19</sup> Nonetheless, most research targets platform molecules with the hope that biorefining can become a sustainable pathway to produce commodity chemicals and fuels. With how established petrochemical-based commodity chemical pathways are in the chemical industry, the production of those compounds from biomass is economically challenged. As a result, higher value-added specialty/fine chemicals could be easier entry markets for biomass-derived products. Within this context, the use of heterogeneous catalysts would allow easy separation and regeneration. One of the most utilized classes of heterogeneous catalysts industrially is zeolites, which are active for a variety of sugar chemistries.<sup>6,8,11,20–25</sup> Interestingly, some of the porous solid acid-catalyzed fructose to HMF studies report observing difructose anhydrides (DFAs) as an intermediate.<sup>26–28</sup> DFAs are higher-value cyclic disaccharides formed from the condensation of two fructose molecules through the formation of two glycosidic linkages. Presently, 15 DFA isomers and their glycosylated derivatives (glycosyl-DFAs) have been identified as having industrial potential in the food, nutraceutical, and pharmaceutical industries. Potential applications range from chemical markers in certain food products, low-calorie sweeteners, to supplements with early studies indicating their potential to prevent the development of cancer, osteoporosis, and anemia. Additionally, their rigidity and hydrophilicity offer possibilities for the development of surfactants, hydrophilic polymers, and complexing agents.<sup>29–31</sup>

Typically, DFAs are currently produced *via* biosynthetic methods with DFA yields of up to 82%,<sup>29</sup> under optimized conditions. However, these processes use expensive tailored DFAs-forming enzymes with costly inulin or levan polysaccharides as substrates.<sup>32</sup> The large-scale production of DFAs using those methods is constrained by the isolation and purification of the main products. In this context, chemocatalytic routes, which employ fructose as a feedstock emerge as an innovative and sustainable catalytic strategy for DFA production. The atom economy of these processes (~90%) is comparable to that of the corresponding biosynthetic processes (~95%) (Table S1), highlighting their potential as efficient and green alternatives.

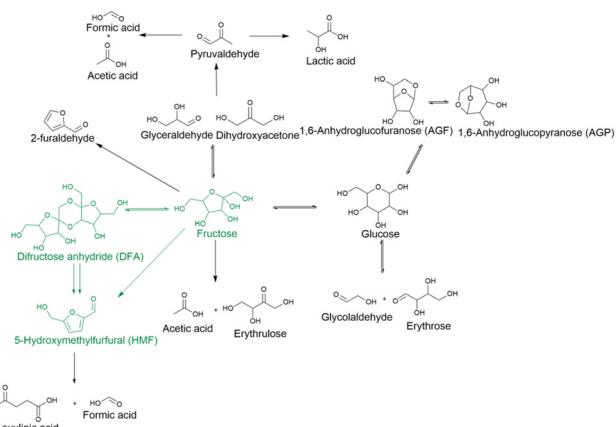
However, prior attempts at producing DFAs through catalytic routes showed significant challenges that include low yields at high conversions (due to low selectivity), or technical issues associated with the use of the hazardous, corrosive, and environmentally unfriendly hydrofluoric acid (HF) or resins with high viscosity, which makes homogenization of the reaction mixture challenging.<sup>33–36</sup> Moreover, most published research focuses on maximizing HMF yields, and little attention was given to DFA production.<sup>27,28,37</sup> The highest DFA yield for a zeolite at a high conversion was 4% for mordenite at 77%

conversion.<sup>28</sup> A sulfonated carbon catalyst in a 1-butyl-3-methyl-imidazolium chloride and 2-methyltetrahydrofuran biphasic system at 100 °C gave a 28.1% DFA yield at 94% conversion.<sup>38</sup> HMF targeted studies usually report relatively high temperatures and comparatively long reaction times, which is likely why DFAs were observed as a short-lived intermediate. Additionally, the catalytic production of furans from carbohydrates in water suffers from low catalytic activity, which increases with the use of polar aprotic solvents, such as gamma-valerolactone (GVL).<sup>39,40</sup> GVL is a biomass-derived, biodegradable solvent, recognized as one of the most promising green alternatives for catalytic biomass transformations.<sup>41,42</sup> However, because sugar solubility in most polar aprotic solvents is typically low at room temperature, water is commonly used in binary solvent mixtures to ensure adequate solubility.<sup>43,44</sup> Solvent selection is moreover important in zeolite-catalyzed biomass conversion processes due to interactions such as competitive adsorption, site-blocking interactions, and various active site restructuring mechanisms. Because of the porous nature of zeolites, molecular sieving of the reaction mixture can take place, partitioning species in or out of the zeolite's pores, which can affect the chemical potential of reactants near the active site.<sup>25</sup> This effect is largely influenced by the size and shape selectivity that the specific dimensions of the zeolite pore openings and channels impose on guest molecules. Studies have also illustrated how sugar molecules can take on different conformations in the confinement of zeolite pores.<sup>22</sup> The solvent system can also organize differently within the zeolite, affording solvent structures that make for easier access to active sites.<sup>45</sup> Nevertheless, the presence of water was particularly noted to be detrimental to some reactions, such as the dehydration of glucose to anhydro-sugars.<sup>46</sup> Therefore, herein, we demonstrate a unique approach for selectively dehydrating fructose to difructose anhydride using acid microporous catalysts in a green polar aprotic solvent. Additionally, we investigated how the presence of water alters the reaction pathway when GVL is used as a solvent, inhibiting DFA formation, an aspect not previously explored in literature, which illustrates how important solvent effects are in these systems. To the best of our knowledge, this is the first time reporting a DFA yield of 75% under mild conditions (1.5 wt% fructose in GVL, 50 mL min<sup>-1</sup> helium flow, 125 °C and 42 μmol of aluminum sites) with a zeolite catalyst. This result indicates that we can achieve a high catalytic performance while maintaining a high standard, resource-efficient, and environmentally responsible approach as reflected by the high reaction mass efficiency comparable or superior to other technologies. (For more information, refer also to Table S1).

## Results and discussion

Fructose conversion is a well-studied process that, depending on the reaction conditions, produces various useful molecules (Scheme 1). Many of those studies showed a solvent depen-





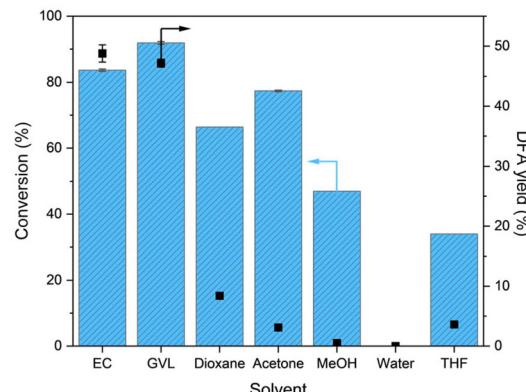
**Scheme 1** Commonly reported fructose conversion pathways. Note – the double arrow indicates the viewpoint of difructose anhydride being an intermediate to 5-hydroxymethylfurfural.

dence on rate and product distribution. These effects stem from competitive adsorption, site-blocking, and intermediate stabilization. Fructose itself can be altered by the solvent system due to the various conformations that the molecule can adopt. These conformations can interact differently with the reaction mixture, *e.g.*, the carbonyl of the keto form (ring-opened form) of fructose coordinating to the Brønsted acid proton. Additionally, from a ring-opened form, the dicyclic DFA molecule is not preferentially formed, but isomerization usually dominates.<sup>8,9,11</sup> Typically, this ring-opened form is the major conformer in mostly aqueous systems, but it is minimal in organic solvents. But even with the ring-closed forms of fructose, the composition of the reaction mixture can impart complications. The way in which fructose can structure with surrounding solvent molecules can hinder its accessibility to active sites. These structuring complications necessitate considering the nature of the solvent utilized for a given reaction due to interactions with both the zeolite framework and the substrate.

### Effect of the solvent on the reaction outcome

To understand the impact of the solvent environment, we conducted fructose dehydration experiments using H-Al $\beta$ 25 (additional details are provided in Fig. S6, S7, S8 and Table S3 in the SI) in several pure solvents that potentially promote the closed-ring form of fructose, all of which are mostly polar aprotic. The polar aprotic solvents employed were tetrahydrofuran (THF), dioxane, ethylene carbonate (EC), acetone, and gamma-valerolactone. For comparison, we used water and methanol (MeOH) as polar protic solvents. With the same reaction conditions (140 °C, 1.5 wt% fructose and 2.4  $\mu$ mol of aluminum sites), we observe that the solvents influence the reaction pathways as fructose conversion and product yields varied significantly (Fig. 1 and Table S3).

In water, which is the co-solvent and a byproduct of the dehydration process, we observed no conversion after 15 min of



**Fig. 1** Conversions and DFA yields for H-Al $\beta$ 25 catalyzed fructose conversion for various solvents at 140 °C, 15 min, 1.5 wt% fructose and 2.4  $\mu$ mol of aluminum sites. Ethylene carbonate, gamma-valerolactone, and acetone have error bars for DFA yield from multiple reactions run.

reaction. This agrees with previous studies indicating low reaction rates for carbohydrate dehydration to furans and anhydrosugars in aqueous environments.<sup>40,46,48,49</sup> Besides, HMF was preferentially produced over 2-furaldehyde (FAL).<sup>48,50</sup> Nonetheless, the formation of humins cannot be discarded, as the solution progressively darkened, and the reaction exhibited poorer carbon balance over time from 100% at 5.8% of conversion after 3 h of reaction to 81% after 24 h and a conversion of 60%.

With THF, the conversion reached 34%, yielding 3.6% DFA, 1.8% HMF, and only 0.3% yield of FAL. In our study, the lack of water initially promoted the production of anhydrous glucose molecules, including 1,6-anhydroglucopyranose (AGP) and 1,6-anhydroglucofuranose (AGF) (8.0% yield), which were generated from the previous isomerization of fructose to glucose (3.2% yield). Notably, the relatively high yield of dihydroxyacetone (4.2%), an intermediate associated with the formation of lactic acid, which typically form under Lewis acidic conditions, suggests that this solvent promoted a different reaction pathway.<sup>51–54</sup> Interestingly, we did not detect formic acid that is often reported as a major product when there is water present. This may be attributed to the low conversion achieved under our conditions. Previous studies using pure THF did not report DFA production.<sup>48,55</sup>

The reaction in methanol produced 47% conversion, with no detection of formic or acetic acid. According to the literature, considerable amounts of formic acid are generated in methanol at high conversion rates, like reactions in water.<sup>48</sup> This result contrasts with our findings at lower conversions. Unidentified species might be intermediates for the formation of esters instead of carboxylic acids, as the latter form when using aqueous solutions of aprotic polar solvents and Lewis acids as catalysts.<sup>54,56</sup>

Furans were the primary products in dioxane with a yield of 19%, while the DFA yield was only 8.4% at a conversion of 66%. As predicted from the literature, the FAL yield was slightly higher than HMF, which might be attributed to the



solvent's low basicity and polarity. At these reaction conditions, we did not detect rehydration of HMF to produce formic acid or levulinic acid, which contrasts with observations reported in the literature. This discrepancy may be explained by the different reaction conditions employed in our study.<sup>44,48</sup>

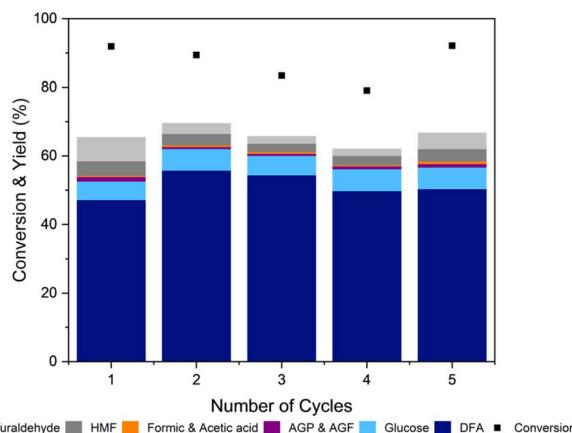
We observed the highest conversions (higher than 77%) at 15 min for acetone, EC, and GVL. However, despite displaying high conversion (77%) in acetone, the DFA yield was low (3.1%). The isomerization pathway to glucose was important, obtaining 16% yield of AGP and AGF in the absence of water. Yet, the overall mass balance was disappointingly low.

EC, a solid at room temperature, shows high viscosity upon melting and toxicity concerns, representing a challenge in handling, scalability, and sustainability compared to other solvents. The conversion of fructose in EC was 84%, with a DFA yield of 49%. Less than 1% of 1,6-anhydroglucose formed despite a 6.9% glucose yield. Formation of furans was also minimal (yields around 3%).

GVL, being a green solvent, offers additional advantages over other organic solvents, such as lower toxicity, improved safety, and better environmental impact, making it a sustainable choice for biomass conversion processes.<sup>57,58</sup> It displayed the highest fructose conversion, 92% after 15 min, with a DFA yield of 47% and a furans yield (HMF and FAL) of 11%.<sup>47</sup> As we increased the reaction time to 3 h, the yield of DFA progressively decreased until it was no longer detected. In contrast, the formation of furans increased to 40%, indicating that DFA further converted into furanic compounds over time.<sup>27,28,33</sup> The conversion of fructose to HMF and DFA proceeded faster than the reverse reaction from DFAs to fructose, as DFAs reversibly retain fructose *via* acid catalysis, delaying its degradation and unexpectedly promoting HMF formation.<sup>27,33,59,60</sup> The FAL yields were consistently higher than the HMF yields, likely due to the high polarity, low basicity of GVL, and limited reactant accessibility, which is in agreement with previous studies.<sup>44,48,49,61</sup> The stability study indicates that the regeneration process retained most of the material's catalytic activity (Fig. 2). This is illustrated by the catalyst crystallinity remaining at 92% *via* XRD (Fig. S13) and the acidity only decreasing by 4% throughout the reuse and regeneration process (Table S2).

Improving DFA yield and production was possible by employing milder reaction conditions (lower temperature, reduced catalyst loading and increasing fructose concentration), preventing further unwanted reactions (Fig. S6, S7 and S8), such as 1,6-anhydroglucose formation as a consequence of fructose isomerization or formic or acetic acid as product degradations.<sup>46,52,60</sup> This result further illustrates the findings of Dumesic *et al.* on the utility of GVL for biomass conversion chemistries.<sup>62</sup>

Evaluation of catalytic performance based on solvent selection shows that GVL and EC are the best solvent choices to target high DFA yields. GVL is the preferred solvent in this case because it is easier to handle compared to EC's solid/waxy nature at lower temperatures.



**Fig. 2** Effect of H-Al<sub>25</sub> zeolite reuse and regeneration on fructose conversion to DFA. Reaction conditions: 1.5 wt% fructose in GVL, 15 min, 140 °C and 2.4 μmol of aluminum sites. Note – for the 1st–4th runs, the catalyst was recovered by washing with water and acetone, followed by drying at 120 °C. For the 5th run, the catalyst was recovered using the same procedure and subsequently calcined at 550 °C.

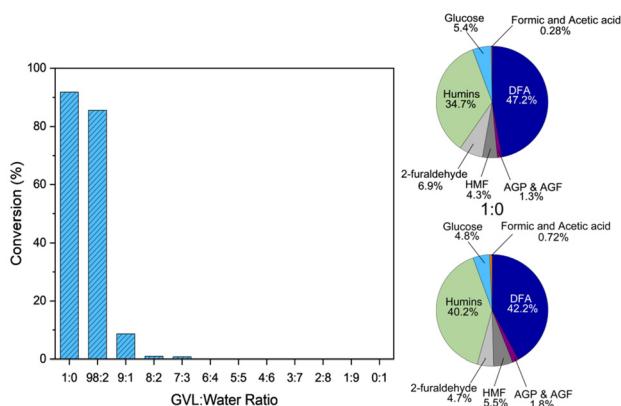
### Effect of water on reactivity

In typical liquid phase reactions, one of the first criteria to evaluate is solvent-reagent miscibility. Thus, many of the reported sugar conversion reactions utilize water-organic mixtures for their solvent systems.<sup>22,63–66</sup> Due to the lack of solubility of fructose in pure GVL, we incrementally added water to the system to increase the solubility of the substrate but also move closer to the same regime that other studies operated at.<sup>25</sup> By doing so, we observed the negative impact of water with a pronounced deactivation of the system at volume percentages of just over 10% (Fig. 3). Beyond the implications of added water on the system, the trace water that GVL will uptake and the water produced during the reaction could also lead to reactivity complications.

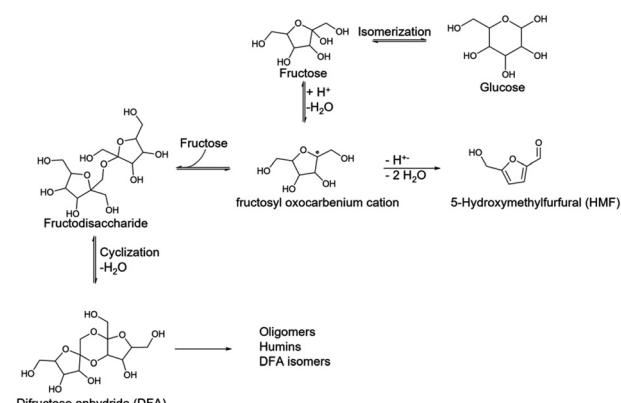
To induce a nearly water-free environment within the constraints of our system set up, we bubbled helium gas (50 mL min<sup>-1</sup>) into the reaction mixture to selectively purge out the water produced during the reaction. This technique aimed to reduce the potential inhibitory effects of ambient and produced water during the reaction on DFA production. In the presence of pure GVL solvent and helium flow, the conversion was 90% after 5 min of reaction and the DFA yield was 75%. This yield remained after an additional 10 min of reaction. In the absence of helium, after 5 min, the conversion was 86%, and the DFA yield was 73%. However, after 15 min, it started to degrade to furans (4.1%) and the DFA yield decreased to 67%. These results may indicate that the water produced during the reaction can lead to the hydrolysis of DFA and push the reaction towards furans. This observation leads us to suggest that the reversibility of DFA production can be impacted by low amounts of water in the reaction mixture, shifting the reaction pathway towards furans (Scheme 2).

Again, leveraging GVL as our preferred solvent, we studied how the incremental addition of water as a cosolvent could





**Fig. 3** Conversion of fructose over H-Al $\beta$ 25 zeolite at 140 °C for 15 min at varying GVL:water ratios. Product yields for 1:0 and 98:2 GVL:water ratio shown in pie charts. Conditions: 3 g solution, 1.5 wt% fructose in GVL, and 2.4  $\mu$ mol of aluminum sites.



**Scheme 2** Proposed mechanism of DFA production from fructose.

impart reaction pathway changes in the system. With only the addition of 2 wt% water to GVL, the conversion decreased from 92% to 86% in the standard 15-minute reaction time. At the 10% water loading condition, the conversion was only 8.7%, with only 1.5% of DFA yield. This decrease in performance illustrates the inhibiting role of water in DFA formation, like the previously reported inhibition with aldehydes and their conversion to anhydrosugars in dioxane.<sup>46</sup> At water percentages of 20% and greater, we observed a complete suppression of the catalytic activity, the mechanism of which is still unclear.

To obtain the kinetic reaction order in fructose and the activation energy of the system, we conducted concentration and temperature dependence experiments (Fig. S9 and S10). At our experimental conditions, we obtained an order of  $1.7 \pm 0.3$  in fructose, and we estimated an activation energy of  $66 \pm 10$  kJ mol<sup>-1</sup>. The obtained kinetic parameters lead us to propose a different reaction pathway rather than DFA being an intermediate to HMF but a competing pathway that involves revers-

ible steps that could lead to HMF (Scheme 2). This proposed pathway requires the activation of fructose at the anomeric carbon to form the fructocarbocation that can undergo dimerization upon contacting another fructose molecule. This dimerized form then undergoes ring-closing dehydration to form DFA, which aligns with the experimentally observed reaction order in fructose.

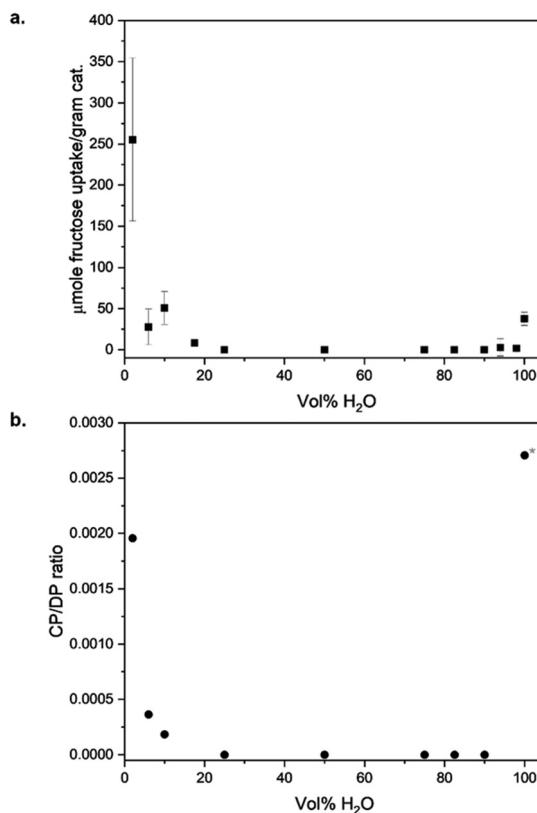
### Effect of water on fructose partitioning

To probe the molecular implications of water addition, we conducted partitioning studies. Given that the reaction rate drops to zero at just over 10% water in the solvent system, understanding the species that can enter the pores of the zeolite at varying water concentrations can provide information about what environment is necessary for catalysis to take place. Utilizing a room temperature mixture of zeolite, GVL, and D<sub>2</sub>O (in place of water for HPLC detection), we quantified the uptake of species into H-Al $\beta$ 25 at various increasing water percentages. We observed that the adsorption of fructose into H-Al $\beta$ 25 only occurred at water concentrations at, or under 10 volume% with decreasing extent of uptake from 2% to 10% water (Fig. 4a). This may be due to water clustering around the Brønsted acid sites within the zeolite,<sup>67</sup> not allowing fructose to enter the porous structure of zeolite beta. This fructose uptake follows the observed water-based deactivation of the catalytic system, which shows no catalytic activity at greater than 10 volume% water content in the solvent system. Such observations further corroborate our hypothesis that the fructose is activated at the Brønsted acid sites within the zeolite pore rather than an outer pore catalytic process.

To further gain molecular insights into the effect of water on this system, we used a solid-state NMR method to gauge the proportion of adsorbed species *versus* outer pore species.<sup>68</sup> Again, utilizing fructose, GVL and water mixtures with increasing water content, we leveraged the favorable polarization transfer between adsorbed species *versus* the lack of polarization transfer between adsorbed species *versus* outer pore/non-confined species must be able to assess the partitioning of these mixtures. The confinement effects that the micropores of the zeolite impart on adsorbed species cause them to have reduced mobility, which allows for residual dipolar couplings between molecules. This preferentially affords proton magnetization transfer to carbon species confined in the pores rather than in a mobile orientation. The more mobile, rapid tumbling nature of outer pore species does not allow for this effect since they have too short relaxation times ( $T_1$ ) to have residual dipolar couplings. Based on this, we can use a cross-polarization pulse sequence to understand the adsorbed species and a direct polarization pulse sequence to observe all the species in the slurry, confined or not, similarly to the work of Martens and coworkers.<sup>68,69</sup> The impact of confinement on the relaxation of species is the major proponent for this magnetization transfer capability.<sup>70,71</sup>

To observe the extent of adsorption of fructose in the H-Al $\beta$ 25 zeolite for the various solvent conditions, we used the ratio of the cross-polarization (absorbed) peak area over the





**Fig. 4** (a) Fructose uptake ( $\mu\text{mol}$  per gram of H-Al $\beta$ 25 zeolite) with increasing water volume percentage in GVL. (b) Anomeric carbon of fructose solid-state Nuclear Magnetic Resonance (ssNMR) cross-polarization peak area and direct polarization Hahn-echo peak area ratio per volume percentage of water in the initial solution the zeolite was contacted with. Note – the asterisk denotes that this value is high due to the ability for the water clusters, likely hydrogen bound to the fructose molecules in the pore, to have a greater cross-polarization effect than in mixtures where GVL is present.

direct excitation (total species) Hahn-echo peak area (Fig. S11). To get the largest impact from the cross-polarization effects coming solely from the residual dipolar couplings of the confined fructose, we used the anomeric carbon chemical shift to track fructose since it does not have a directly bound hydrogen that can provide cross-polarization. This experiment yielded a similar exponential decay as the uptake data (Fig. 4b) and aligns well with the deactivation trend from reactivity data (Fig. 2). These data seem to suggest that the uptake of fructose is necessary for the conversion of fructose to DFA, and outer pore catalysis is not the dominant pathway for this system. It is important to note that the 100 vol% water condition shows the most adsorbed fructose, likely due to the cross-polarization enhancement from the surrounding water molecules solvating the fructose molecules in the pore of the zeolite, since there is no water-organic solvent competition for solvating fructose. The higher availability of protons to donate to the anomeric carbon of fructose in a water rich environment within the zeolite pores is the dominant mode for this signal increase rather than any relaxation effects.

## Conclusions

In this work, we report on a novel catalytic process for the conversion of fructose to DFA in appreciable yields and provide a molecular understanding of the impact of water as a co-solvent on the uptake of fructose into the porous structure of zeolite beta. We demonstrated that H-Al $\beta$ 25 is an active catalyst and that biomass-derived GVL is the preferential solvent to drive this reaction pathway towards DFAs. When utilizing incremental concentrations of water as a co-solvent to increase fructose solubility, we observed an inhibition effect with greater than 10 vol% water. Partitioning studies and ssNMR show that at these increasing water concentrations, fructose can no longer enter the pores of zeolite beta, leading to a decrease in fructose conversion. Through the spectroscopic and chromatographic observations obtained in this work, we suggest that intraporous Brønsted acid site accessibility is pivotal to produce DFA through heterogeneously catalyzed routes.

## Author contributions

I. H-S. and F. R-R. performed reactivity testing and analysis. F. I. performed ssNMR studies and analysis. F. I. and G. G. performed partitioning studies and analysis. E. A. L-R., I. H-S., F. I., R. M., and F. R-R. performed catalyst characterization. I. H-S., F. I., E. A. L-R., I. H., and N. C-M. conceptualized the project. The manuscript was written by I. H-S., and F. I. and revised by E. A. L-R., M. A., I. H., and N. C-M. with input from all authors. N. C-M., and I. H. were responsible for project administration.

## Conflicts of interest

There are no conflicts to declare.

## Data availability

The authors declare that all the data supporting the findings of this study are available within the article (and supplementary information (SI) files), or available from the corresponding author on reasonable request.

Supplementary information is available. See DOI: <https://doi.org/10.1039/d5gc04714e>.

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