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## ARTICLE

Synthesis of heteroleptic calcium amide complexes *via* manipulation of the Schlenk equilibriumMatthew de Vere-Tucker,<sup>ab</sup> Alexandra M. Z. Slawin<sup>b</sup>, David B. Cordes<sup>b</sup> and Andreas Stasch<sup>\*b</sup>Received 00th January 20xx,  
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Heteroleptic calcium  $\beta$ -diketiminato (BDI) complexes have been recognised as potent (pre)catalysts for a variety of transformations and have demonstrated valuable stoichiometric reactivity. Bulky ligands such as BDIs have been utilised to prevent unwanted dimerisation and decomposition reactions of active heteroleptic species to their homoleptic congeners. Homoleptic bis-ligated alkaline earth metal (Ae) species of the type "L<sub>2</sub>Ae" have hitherto been often considered kinetic sinks and less valuable byproducts. This work shows that through the use of the sterically demanding prolignand HC(*i*PrCNDip)<sub>2</sub>H, Dip = 2,6-*i*Pr<sub>2</sub>C<sub>6</sub>H<sub>3</sub> (<sup>*i*Pr</sup>DipBDIH), the homoleptic calcium complex [(<sup>*i*Pr</sup>DipBDI)<sub>2</sub>Ca] can be destabilised to the degree that when treated with calcium bis(hexamethyldisilazide) (Ca(HMDS)<sub>2</sub>), a thermal dimerisation of the ligands takes place to afford the heteroleptic complex [(<sup>*i*Pr</sup>DipBDI)Ca(HMDS)]. DFT calculations support that this reaction is enthalpically unfavourable, an effect that is reduced by the destabilising influence of the bulk in [(<sup>*i*Pr</sup>DipBDI)<sub>2</sub>Ca], but entropically made possible. This synthesis has been extended to a convenient one-pot regime from [(<sup>*i*Pr</sup>DipBDI)K], CaI<sub>2</sub> and Ca(HMDS)<sub>2</sub>, and the resultant calcium amide complex has been utilised as a precursor to prepare a thermally stable, donor-solvent-free molecular calcium hydride complex [(<sup>*i*Pr</sup>DipBDI)CaH]<sub>2</sub>.

## 1. Introduction

Heteroleptic complexes of alkaline earth (Ae) metals are strongly prone to Schlenk-type redistribution reactions.<sup>1,2</sup> These were first described by Schlenk and Schlenk Jr in 1929 as it was observed that an ethereal solution of a Grignard reagent contained not only the RMgX species (where R is an alkyl or aryl substituent and X is a halide), but also the R<sub>2</sub>Mg and MgX<sub>2</sub> species.<sup>3</sup> The Schlenk equilibrium (Fig. 1) manifests itself due to the predominant ionic nature of the bonds between the metal ion and the coordinated ligands.<sup>4</sup> These are generally easy to distort, which allows the relatively facile formation and breaking of coordination bonds. As the metal ion size increases down a group, the propensity towards Schlenk-type redistribution increases, leading to faster ligand exchange rates and often greater instability of heteroleptic complexes. Furthermore, purely inorganic AeX<sub>2</sub>-type species have high lattice energies, which can further drive Schlenk-type redistribution to one side. This is due to the precipitation of the AeX<sub>2</sub> species from solution and thus the equilibrium is driven towards the redistributed species. Related to the use of Ae metal complexes in catalysis, active heteroleptic complexes can be generated *in-situ* from homoleptic precatalyst complexes and are then subject to these equilibria.<sup>1,2</sup>

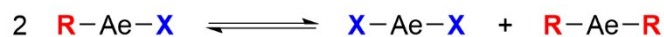


Fig. 1 Simplified Schlenk-type equilibrium

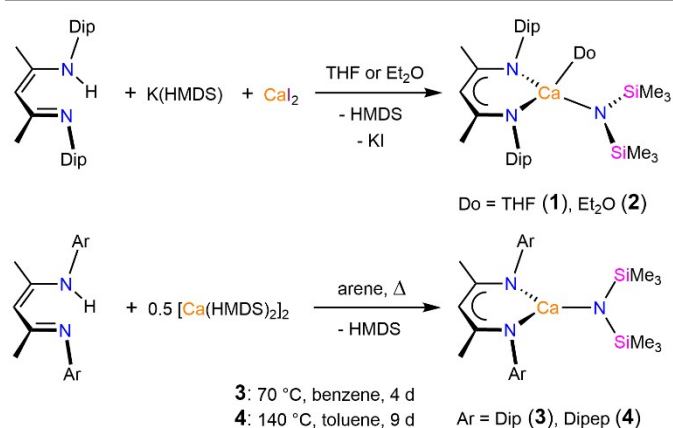
In addition, when LAeX species (where L is a sterically demanding anionic spectator ligand) dimerise, the resultant L<sub>2</sub>Ae species are generally considered to be kinetically inert. As such, both the often valuable ligand entity, L, and the reactive fragment, X, e.g. a metal hydride moiety, can be lost for the desired reactivity, for example, as part of catalytic processes. This redistribution can be prevented by kinetically suppressing exchange processes, or by manipulating the position of the equilibrium through donor-solvent effects.<sup>5,6</sup>  $\beta$ -Diketiminato (BDI) ligands are ubiquitous *N,N'*-chelating ligands in main group chemistry and are particularly widely used in Ae metal chemistry for their strong stabilising characteristic.<sup>7–13</sup> BDI ligand-stabilised calcium complexes have applications in catalysis,<sup>14–20</sup> and recently have been used as precursors for low-valent synthons.<sup>21–25</sup> Calcium hydride complexes are a particularly rich field of exploration, with molecular calcium hydride species demonstrating rich and potent stoichiometric and catalytic reactivity.<sup>26–34</sup>

A key precursor to the synthesis of calcium hydride complexes is typically a calcium HMDS species (HMDS = hexamethyldisilazide, N(SiMe<sub>3</sub>)<sub>2</sub>). For BDI ligand-stabilised calcium HMDS complexes, there are two synthetic routes available, see Scheme 1. One approach is the treatment of the prolignand with KHMDS to effect the *in situ* deprotonation and subsequent salt metathesis with CaI<sub>2</sub> in an ethereal solvent<sup>35,36</sup> to afford donor solvent adducts such as [(<sup>*Me*</sup>DipBDI)Ca(HMDS)(THF)] **1** and [(<sup>*Me*</sup>DipBDI)Ca(HMDS)(OEt<sub>2</sub>)] **2**

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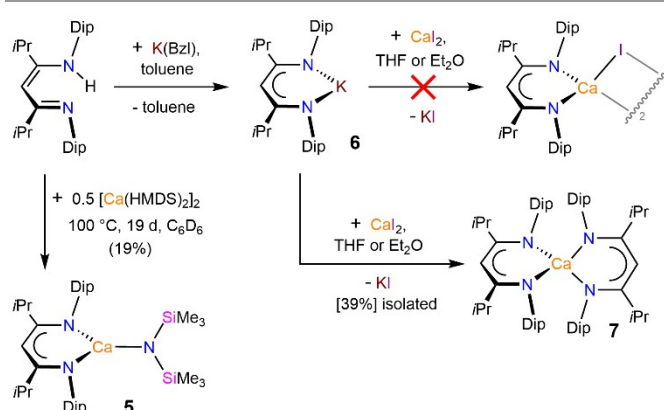
(MeDipBDI = (HC(MeCNDip)<sub>2</sub>; Dip = 2,6-diisopropylphenyl). Another route utilises direct deprotonation of a suitable BDIH prolignand with Ca(HMDS)<sub>2</sub> to afford complexes such as [(MeDipBDI)Ca(HMDS)] **3** and [(MeDipecBDI)Ca(HMDS)] **4** (MeDipBDI = (HC(MeCNDip)<sub>2</sub>; Dipec = 2,6-bis(3-pentyl)phenyl)).<sup>37</sup> The latter route, while slower, has the advantage of avoiding coordinating solvents and halide salts.

We have recently developed a facile route to the synthesis of the isopropyl-substituted prolignand *i*PrDipBDIH<sup>38,39</sup> (= (HC(*i*PrCNDip)<sub>2</sub>H)) which we have employed for the stabilisation of a range of highly reactive magnesium species that highlighted the impact and importance of backbone substituent modifications.<sup>39,40</sup> He were demonstrate these substituent effects on the synthetic accessibility and stability of heteroleptic calcium amide and hydride complexes.

## 2. Results and discussion

It was presumed that the treatment of *i*PrDipBDIH with Ca(HMDS)<sub>2</sub> would yield the desired BDI calcium amide precursor complex [(*i*PrDipBDI)Ca(HMDS)] **5** via an elimination reaction as shown in Scheme 1. However, no significant reaction was observed *via* <sup>1</sup>H NMR spectroscopy even when a solution of *i*PrDipBDIH and Ca(HMDS)<sub>2</sub> was heated in benzene-*d*<sub>6</sub> to 100 °C for over one week. This resistance to deprotonation was surprising, as MeDipBDIH is readily deprotonated through its treatment with Ca(HMDS)<sub>2</sub>.<sup>37</sup> After nineteen days at 100 °C, resonances presumably corresponding to the formation of complex **5** were observed (*vide infra*); however, only with a conversion of ca. 17%, and thus an alternative synthetic route was targeted. It should be noted that this extremely slow reaction is proposed to be likely a result of kinetic factors, as previously shown by other successful attempts to deprotonate *i*PrDipBDIH with alkali metal amide bases.<sup>38</sup> It was postulated that the deprotonation was the rate-limiting step and as such a route where *i*PrDipBDIH is deprotonated with a more potent base was necessary. Thus, *i*PrDipBDIH was treated with a slight excess of benzyl potassium, K(Bzl), in deuterated benzene, which, after one hour at room temperature, afforded a red-orange solution (Scheme 2). <sup>1</sup>H NMR spectroscopic analysis indicated the formation of the potassium complex [(*i*PrDipBDI)K] **6**, per the

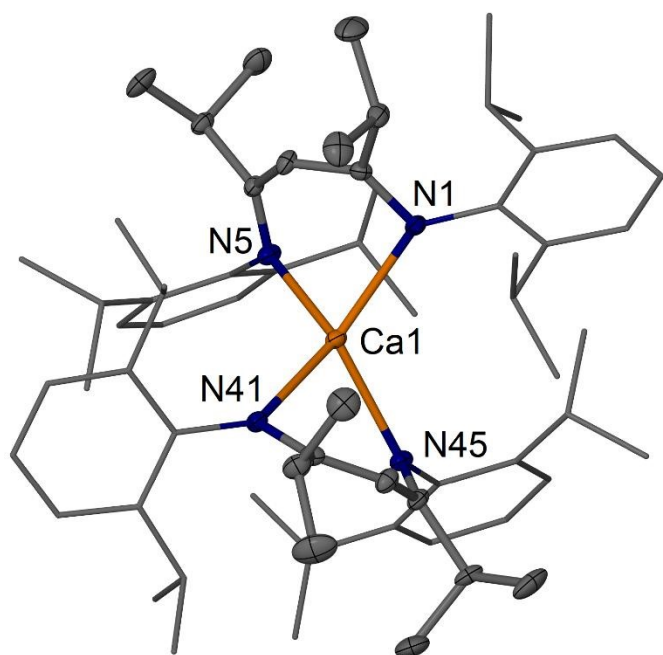
disappearance of the singlet at  $\delta = 12.30$  ppm corresponding to the imino proton, the appearance of a resonance for the by-product, toluene), and the characteristic backbone methine resonance at 4.16 ppm.<sup>38</sup>



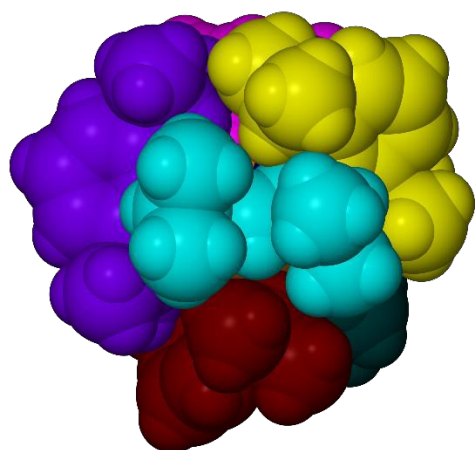
With the intention of obtaining the heteroleptic calcium iodide species [(*i*PrDipBDI)CaI]<sub>2</sub>, complex **6** was treated with CaI<sub>2</sub> under a variety of conditions, for example heating in an arene solvent, stirring in Et<sub>2</sub>O at room temperature, 0 °C and -78 °C or stirring the mixture in THF at room temperature and -78 °C, etc, but a heteroleptic complex of this type was never obtained in any observable quantities. Instead, the only main products obtained were either the prolignand *i*PrDipBDIH or the homoleptic complex [(*i*PrDipBDI)<sub>2</sub>Ca] **7**, see Scheme 2. This salt-metathesis route has previously been used with some success for the synthesis of heteroleptic BDI calcium iodide species with methyl backbone BDI ligands<sup>21,41</sup> and it is notable that the homoleptic complex [(MeDipBDI)<sub>2</sub>Ca] has previously been reported as a by-product for this process.<sup>41</sup>

The <sup>1</sup>H NMR spectrum of [(*i*PrDipBDI)<sub>2</sub>Ca] **7** is worthy of comment and is consistent with NMR spectra of other bulky homoleptic alkaline earth metal BDI species.<sup>42</sup> The high degree of steric crowding of complex **7** is evidenced by the splitting of the isopropyl methyl and methine resonances resulting into six septets and twelve doublets, corresponding to the methine and methyl groups, respectively, rather than the usual three doublets and two septets as would be expected for *i*PrDipBDI-stabilised complexes, where the solution-averaged structure shows one *i*PrDipBDIM environment with equivalence above and below the NCCCN plane and across the left- and right-hand sides of the ligand moiety. This highly asymmetric ligand environment in the solution state is due to the crowded and interlocked nature of the complex. There is a substantial twist in the NCCCN plane observed in the solid-state structure (*vide infra*) which appears to be retained in solution and thus results in the unusual separation of the backbone isopropyl resonances. The two ligands themselves are equivalent according to NMR spectroscopy, as evidenced by the single backbone methine resonance. Storage of a saturated diethyl ether solution of complex **7** for several days at -40 °C afforded single crystals suitable for X-ray crystallography, see Fig. 2 and 3. Complex





**Fig. 2** Molecular structure of  $[(iPrDipBdi)_2Ca] \cdot OEt_2$ , **7**· $OEt_2$  (thermal ellipsoids with 30% probability, hydrogen and solvent atoms omitted for clarity). The Dip-substituents are shown as wireframe only. Selected bond lengths (Å) and angles (°): Ca1–N1 2.383(2), Ca1–N5 2.412(2), Ca1–N41 2.389(3), Ca1–N45 2.418(2); N1–Ca1–N5 85.13(8), N41–Ca1–N45 85.13(8), N1–Ca1–N41 132.06(8), N1–Ca1–N45 118.49(8), N41–Ca1–N5 118.20(9), N5–Ca1–N45 122.66(9).

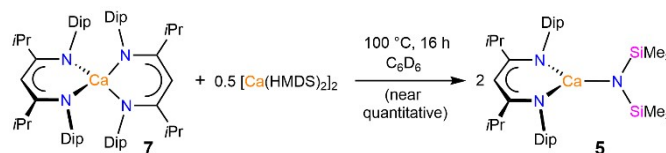


**Fig. 3** Space-filling (van der Waals) model of the molecular structure of  $[(iPrDipBdi)_2Ca] \cdot OEt_2$ , **7**· $OEt_2$ , without the solvent molecule. The  $iPrCC(H)CIPr$  backbone units are turquoise (on N41, N45) and dark green (on N1, N5), and the Dip substituents are yellow (on N41), purple (on N45), dark red (on N1) and magenta (on N5); Ca is orange and N atoms are blue but are hardly visible in this view.

$[(iPrDipBdi)_2Ca] \cdot OEt_2$ , **7**· $OEt_2$ , crystallised in the triclinic space group  $P\bar{1}$  with a full molecule and one equivalent of diethyl ether in the asymmetric unit. The calcium atom is coordinated in a distorted tetrahedral fashion by the two  $N,N$ -chelating BDI ligands, which are approximately orthogonal to one another with respect to their  $NCaN$  least-square planes ( $82.2(9)^\circ$ ). The significant steric strain of this complex is exemplified by the displacement of the Ca from the two  $NCCCN$  least square planes, 1.497(3) and 1.522(3) Å, respectively. These values are greater than those of  $[(MeDipBdi)_2Ca]$ <sup>35,42,43</sup> and notably, given the smaller ionic radius of calcium compared to strontium and barium<sup>44</sup>, larger than that of the strontium and barium

complexes  $[(MeDipBdi)_2Ae]$  ( $Ae = Sr, Ba$ ) ( $Ca = 1.291$  Å,  $Sr = 1.381$  Å and  $Ba = 1.498, 0.946$  Å). The Ca–N bond distances range from 2.383(2) to 2.418(2) Å, which are slightly larger than those of  $[(MeDipBdi)_2Ca]$  (2.374 and 2.384 Å). The coordination environment around Ca, as judged by the  $\tau_4$  values, is slightly less tetrahedral for **7** (0.74) compared with  $[(MeDipBdi)_2Ca]$  (0.78). The steric strain and interlocking of the substituents in complex **7** can be visualised in the space-filling model, Fig. 3, which also points to the substantial stabilisation from dispersion between substituents. The twisting and distortion of the  $iPrCC(H)CIPr$  backbone units, e.g. the central turquoise unit shown in Fig. 3, are due to the steric congestion, but it is worth pointing out that the BDI backbone methine C(H) unit (centre of Fig. 3) that can act as a coordination site, is visible.

The apparent substantial steric strain in complex **7** prompted speculation that the  $iPrDipBdi$  ligands may engage in ligand exchange reactions at high temperatures. As such, a deuterated benzene solution of **7** was treated with one equivalent of  $Ca(HMDS)_2$  and the mixture was heated to 100 °C for 12 hours.  $^1H$  NMR spectra of the mixture after this period show the resolution of the six septets and twelve doublets down to resonances for a new compound with two septets and three doublets, indicating the full consumption of complex **7** and suggesting the formation of a sterically less congested species, see Scheme 3. Furthermore, loss of the resonance at  $\delta = 0.32$  ppm corresponding to the  $Ca(HMDS)_2$  hydrogens and the appearance of a resonance at  $\delta = 0.12$  ppm (consistent with that observed for  $[(MeDipBdi)Ca(HMDS)]$ <sup>36</sup>), indicate the formation of  $[(iPrDipBdi)Ca(HMDS)]$  **5**. Storage of a saturated  $n$ -hexane solution of **5** at  $-40$  °C afforded large crystals of the donor solvent-free complex  $[(iPrDipBdi)Ca(HMDS)]$  **5**, see Fig. 4, whereas storage of a concentrated diethyl ether solution of **5** afforded large crystals of the diethyl ether adduct  $[(iPrDipBdi)Ca(HMDS)(OEt_2)]$  **8**, Fig. 5.

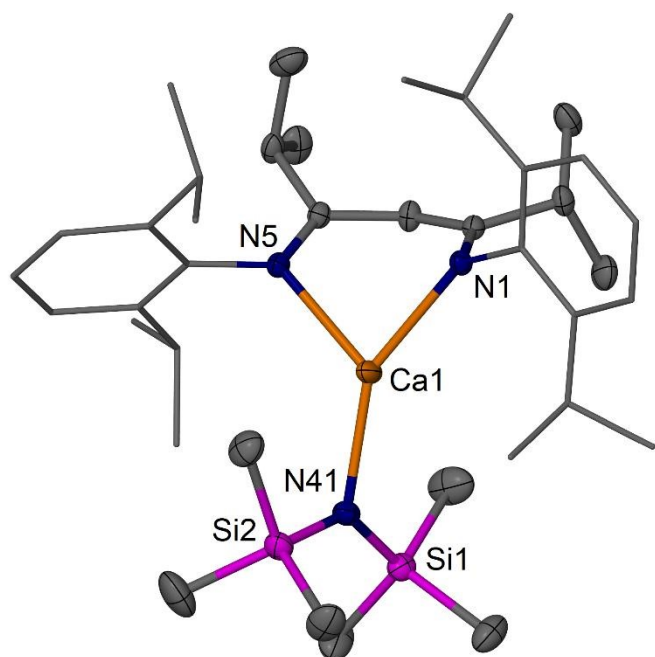


**Scheme 3** Synthesis of  $[(iPrDipBdi)Ca(HMDS)]$  **5**.

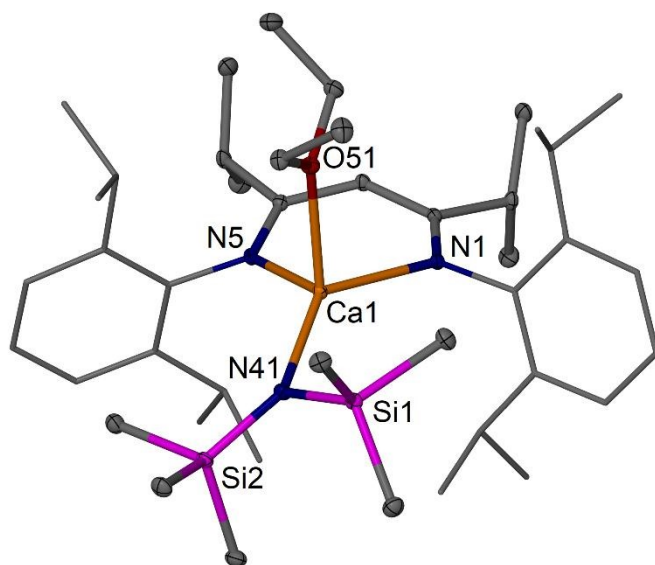
**Compounds**  $[(iPrDipBdi)Ca(HMDS)]$  **5**,  $[(iPrDipBdi)Ca(OEt_2)(HMDS)]_2$  **8** and the highly similar solvate  $[(iPrDipBdi)Ca(OEt_2)(HMDS)] \cdot OEt_2$ , **8**· $OEt_2$ , see Fig. S15, all crystallised with one main molecule in the asymmetric unit in a monoclinic crystal system. Comparing the molecular structures of complexes **5**, **8** and  $[(iPrDipBdi)_2Ca]$  **7** with their  $MeDipBdi$  counterparts (**3**, **2** and  $[(MeDipBdi)_2Ca]$ ) revealed highly comparable Ca–N bond lengths, with marginally longer bonds observed for the solvated species, as would be expected. The bulkier backbone units in **5** and **8** appear slightly more twisted compared with their  $MeDipBdi$ -analogues. The N1–Ca1–N5 bond angle in **5** ( $84.87(4)^\circ$ ) is somewhat wider than that of **8** ( $82.07(3)^\circ$ ) and **3** ( $82.11(11)^\circ$ ) and both are slightly narrower than those of **7** ( $85.13(9)$  and  $85.10(9)^\circ$ ). The sum of the three N–Ca–N bond angles in **5** ( $360.0(1)^\circ$ ) shows planarity, whereas







**Fig. 4** Molecular structure of compound  $[(iPrDipBDI)Ca(HMDS)]$  **5** (thermal ellipsoids with 30% probability, hydrogen and solvent atoms omitted for clarity). The Dip-substituents are shown as wireframe only. Selected bond lengths (Å) and angles (°): Ca1–N1 2.3131(12), Ca1–N5 2.3128(13), Ca1–N41 2.2631(13), Si1–N41 1.6804(14), Si2–N41 1.6796(14); N5–Ca1–N1 84.87(4), N41–Ca1–N1 147.55(5), N41–Ca1–N5 127.55(5), Si1–N41–Ca1 116.77(7), Si2–N41–Ca1 111.60(7), Si2–N41–Si1 131.55(8).

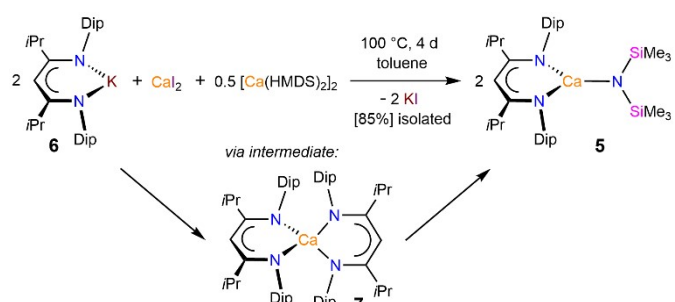


**Fig. 5** Molecular structure of  $[(iPrDipBDI)Ca(HMDS)(OEt_2)]$  **8** (thermal ellipsoids with 30% probability, hydrogen and solvent atoms omitted for clarity). The Dip-substituents are shown as wireframe only. Selected bond lengths (Å) and angles (°): Ca1–N1 2.4239(9), Ca1–N5 2.3802(9), Ca1–N41 2.3435(9), Ca1–O51 2.3810(8), Si1–N41 1.7096(10), Si2–N41 1.7173(9); N5–Ca1–N1 82.07(3), N41–Ca1–N1 136.49(3), N41–Ca1–N5 126.61(3), N41–Ca1–O51 106.57(3), Si1–N41–Si2 118.67(5), Si2–N41–Ca1 130.02(5), Si1–N41–Ca1 110.81(4).

that of **8** ( $345.2(1)^\circ$ ) is closer to planar than tetrahedral geometry. The most dramatic difference between **5** and **8** is in the magnitude of the deviation of Ca from the NCCCN least-square plane, which for **5** is  $1.5145(16)$  Å, which is comparable with that of **7** and near double that of **8** ( $0.7588(14)$  Å), albeit

more consistent with that of **3** ( $1.333$  Å). Also notable is the significant difference in the dihedral angle of the Ca–N51 unit in **8**, which at  $28.1(3)^\circ$  is approaching coplanarity with the  $\beta$ -diketiminato moiety due to the  $Et_2O$  ligand, whereas for **5** it is nearly orthogonal at  $85.38(10)^\circ$ . Unlike in the packing of  $[(MeDipBDI)Ca(HMDS)]$  **3**, only weak intramolecular interactions are observed in **5**, including  $Ca \cdots H(iPr)$  (shortest:  $2.7550(5)$  Å) and  $Ca \cdots H(Me_3SiN)$  (shortest:  $2.8124(6)$  Å), and Si–N–Ca angles of  $111.60(7)$  and  $116.77(7)^\circ$ . The buried volumes ( $V_{bur}$ )<sup>45,46</sup> of the  $iPrDipBDI$  ligands were calculated to be 49.4% for **5** and 45.5% for **8** – perhaps counterintuitively, but as the Ca centre deviates from the ligand plane in **5** it is brought closer to the isopropyl groups. The calculated  $V_{bur}$  for **5** is notably greater than that of **3** (46.9%).

Due to the low overall isolated yield of complex **7** (39% isolated over three crystalline crops) from the reaction of  $[(iPrDipBDI)K]$  **6** with  $CaI_2$  in a 1:1 molar ratio at  $-78^\circ C$  in  $Et_2O$ , a more convenient synthetic route was sought. A one-pot approach was tested by heating a toluene slurry of  $[(iPrDipBDI)K]$  **6**, finely ground  $CaI_2$  and  $Ca(HMDS)_2$  to  $100^\circ C$  for 4 days with rapid stirring, see Scheme 4. This afforded complex **5** as a pale brown solid in 85% yield after workup. It was hoped that if it was possible to perform this redistribution reaction with sparingly soluble  $CaI_2$  in an arene solvent, then perhaps it could be possible to directly treat the homoleptic complex **7** with commercial  $(CaH_2)_\infty$  to afford the heteroleptic calcium hydride complex  $[(iPrDipBDI)CaH_2]$  **9**. As such, a deuterated benzene solution of **7** was treated with 10 equivalents of very finely ground  $(CaH_2)_\infty$ . No reaction was observed, however, even after heating the mixture to  $100^\circ C$  for seven days. The effective complete insolubility of  $CaH_2$  is likely a significant factor here, coupled with its high lattice energy ( $-2401$  kJ mol<sup>-1</sup>) that is substantially greater than that of  $CaI_2$  ( $-1905$  kJ mol<sup>-1</sup>).<sup>47</sup>

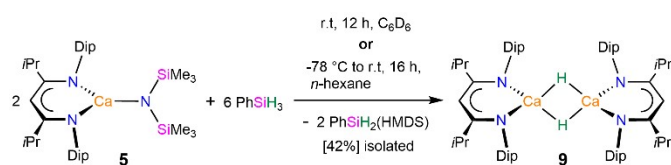


**Scheme 4** One-pot synthesis of  $[(iPrDipBDI)Ca(HMDS)]$  **5**.

Molecular main group hydrides<sup>34,48,49</sup> are of interest due to their intrinsic high reactivity and their range of applications, from catalysis, to energy storage and materials deposition. Calcium hydride complexes are generally synthesised *via* ligand metathesis of the HMDS moiety with a hydride source such as phenylsilane.<sup>21,26,50,51</sup> As such, a deuterated benzene solution of complex **5** was treated with three equivalents of phenylsilane, see Scheme 5. Analysis *via*  $^1H$  NMR spectroscopy after 12 hours showed the complete disappearance of the peaks corresponding to **5**, and the formation of a new set of  $iPrDipBDI$  resonances. Singlets are observed at  $\delta = 0.22$  ppm, integrating



to 18 protons, and  $\delta = 5.16$  ppm, integrating to two protons. These are consistent with the formation of  $\text{PhSiH}_2\text{N}(\text{SiMe}_3)_2$ . Furthermore, an apparent new BDI backbone methine resonance at the more upfield chemical shift of  $\delta = 4.79$  ppm is found, as is a broad singlet at  $\delta = 4.29$  ppm, which is consistent with that of a hydride moiety, as observed for previously isolated BDI calcium hydride species.<sup>26,50,51</sup>  $^1\text{H}$ - $^{13}\text{C}$  and  $^1\text{H}$ - $^1\text{H}$  2D NMR experiments confirmed that the singlet at  $\delta = 4.29$  ppm shows no cross-peaks with any other resonances, as would be expected for a metal hydride resonance, indicating the synthesis of  $[(\text{iPrDipBDI})\text{CaH}]_2$  **9**. To obtain isolated material, an *n*-hexane solution of three equivalents of phenylsilane was added dropwise to a stirred *n*-hexane solution of **5** at  $-78^\circ\text{C}$ . This afforded a colourless solid and pale green solution. The colourless solid was found to be deuterated benzene soluble and  $^1\text{H}$  NMR spectroscopic analysis revealed it to be pure complex **9**, afforded in a 42% yield. The donor adduct  $[(\text{iPrDipBDI})\text{CaH}(\text{OEt}_2)_2]$ , c.f.  $[(\text{MeDipBDI})\text{CaH}(\text{OEt}_2)_2]$ <sup>52</sup>, was generated *in-situ* from  $[(\text{iPrDipBDI})\text{Ca}(\text{HMDS})(\text{OEt}_2)]$  **8** in a similar manner but was not isolated.



Scheme 5 Synthesis of  $[(\text{iPrDipBDI})\text{CaH}]_2$  **9**.

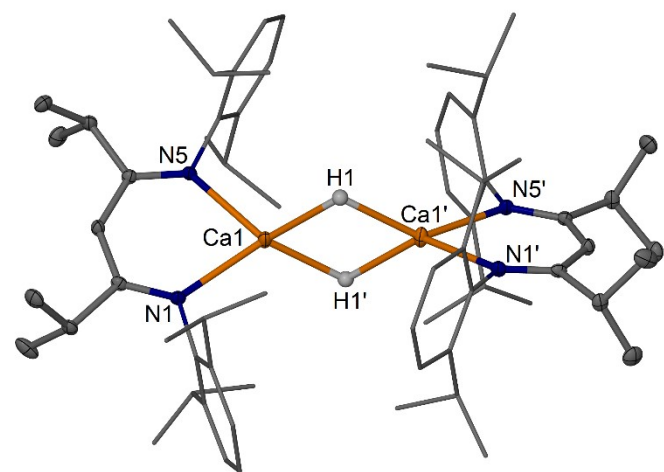


Fig. 6 Molecular structure of  $[(\text{iPrDipBDI})\text{CaH}]_2$  **9** (ellipsoids with 30% probability, hydrogen atoms except for H1 and H1' are omitted). The Dip-substituents are shown as wireframe only. Selected bond lengths (Å) and angles ( $^\circ$ ): Ca1–N1 2.3324(17), Ca1–N5 2.3316(17), Ca1–H1 2.103(7), Ca1–H1' 2.22(3), Ca1...Ca1' 3.4567(9), N5–Ca1–N1 79.75(6), N1–Ca1...Ca1' 142.79(4), N5–Ca1...Ca1' 137.23(4).

Crystals of  $[(\text{iPrDipBDI})\text{CaH}]_2$  **9** suitable for X-ray diffraction were obtained from the storage of the concentrated supernatant solution. Compound **9** crystallised in the tetragonal space group  $I4_1cd$  with half a molecule in the asymmetric unit, see Fig. 6. The crystal structure confirms the proposed dimeric nature of the complex. The mean Ca–N bond length (2.332 Å) of **9** is slightly longer than that of the previously reported  $[(\text{MeDipBDI})\text{CaH}]_2$  **10** (2.316 Å). The Ca1...Ca1' distance in **9** is around 0.1 Å larger, at 3.4566(6) Å versus 3.3401(6) Å in **10**. There is negligible

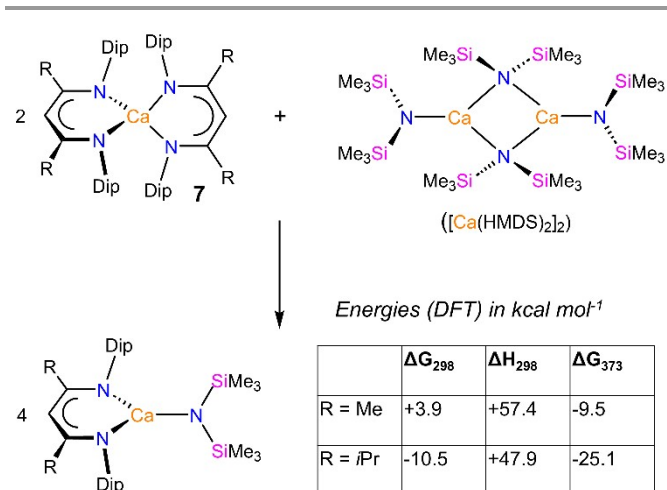
deviation (0.182(3) Å) of the Ca from the NCCCN least-square plane (c.f. 0.437 Å for **10**) and the amide precursor **5** (1.5145(16) Å, *vide supra*). The small hydride ligands in **9** and the relatively planar NCCNCa-ring also lead to the least twisting of the backbone unit in the characterised series. The dihedral angle about the Ca...Ca axis in **9** (39.69(7) $^\circ$ ) is substantially different to the co-planar arrangement of the two  $\text{MeDipBDI}$ Ca moieties in **10**. The  $V_{\text{bur}}$  of complex **9** has been determined as 52.4%, which is notably greater than the  $V_{\text{bur}}$  of **10** (49.5%), and surprisingly, even similar or slightly greater than that of  $[(\text{DipBDI})\text{CaH}]_2$  **11**<sup>21</sup> (51.8%). This may indicate that the backbone *iPr* groups provide notable additional steric protection, although care should be taken when closely comparing these numbers as packing effects will influence these values. This increased  $V_{\text{bur}}$  appears not to be simply a function of the isopropyl groups tightening the ligand bite angle with an N–Ca–N angle of 79.75(6) $^\circ$  for **9**, which is essentially identical to that in **11** (79.25(5) $^\circ$ ). Comparing the Ca–N–C<sub>ipso</sub> angles, **9** exhibits a slightly tighter to identical angle (108.87(12) $^\circ$  for Ca–N1–C<sub>ipso</sub>, 108.29(12) $^\circ$  for Ca–N5–C<sub>ipso</sub>) compared to those of **11** (109.48(9) $^\circ$  and 110.5(1) $^\circ$ ). The steric bulk in **9** is significantly more symmetrically distributed due to the lower steric demand of the hydride ligands compared to the HMDS groups in  $[(\text{iPrDipBDI})\text{Ca}(\text{HMDS})]$  **5** and  $[(\text{iPrDipBDI})\text{Ca}(\text{HMDS})(\text{OEt}_2)]$  **8**, see Fig. S16, S17.

Preliminary investigations towards the thermal stability of  $[(\text{iPrDipBDI})\text{CaH}]_2$  **9** have been carried out, with a deuterated benzene solution of **9** being heated to 100  $^\circ\text{C}$  with periodic  $^1\text{H}$  NMR analysis to determine the rate of decomposition. It was found that **9** has a half-life of approximately 96 hours in benzene at 100  $^\circ\text{C}$  (Fig. S13). The stability of  $[(\text{MeDipBDI})\text{CaH}]_2$  **10** was probed in addition under the same conditions (Fig. S14), with 50% decomposition largely to  $[(\text{MeDipBDI})_2\text{Ca}]$  observed after ca. 30 hours, providing further evidence for the increased complex robustness afforded by the BDI backbone *isopropyl* groups of **9** compared to the methyl groups of **10**.

For further insights into this apparent reversion of the conventional direction of the Schlenk-type equilibrium towards the formation of the heteroleptic complex  $[(\text{iPrDipBDI})\text{Ca}(\text{HMDS})]$  **5**, the Gibbs Free Energies for the redistribution reaction of  $[(\text{iPrDipBDI})_2\text{Ca}]$  **7** and  $\text{Ca}(\text{HMDS})_2$ , modelled as the unsolvated dimer  $[\text{Ca}(\text{HMDS})_2]_2$ , were investigated by DFT calculations at the B3LYP<sub>PCM(benzene)</sub>/6-311G\*\*//B3LYP/6-31G\*\* level of theory, and compared to the same reaction for the  $\text{MeDipBDI}$  ligand, see Fig. 7.

For both cases, the reaction requires a significant investment in reaction enthalpy but becomes viable for entropic reasons, especially at higher temperatures. It was found that in the case of  $[(\text{MeDipBDI})_2\text{Ca}]$  plus  $\text{Ca}(\text{HMDS})_2$ , the formation of the heteroleptic complex  $[(\text{MeDipBDI})\text{Ca}(\text{HMDS})]$  is approximately energy neutral across the studied temperature range, whereas the formation of  $[(\text{iPrDipBDI})\text{Ca}(\text{HMDS})]$  **5** from ligand redistribution is exergonic ( $-25.1$  kcal mol $^{-1}$  for the shown reaction, or  $-6.3$  kcal mol $^{-1}$  as referenced to the formation of one mole of **5**) at the reaction temperature of 100  $^\circ\text{C}$  for the ligand with the bulkier backbone. The high reaction temperature is likely also required for kinetic reasons. The DFT





**Fig. 7** DFT-calculated energies (B3LYP<sub>PCM(benzene)</sub>/6-311G\*\*//B3LYP/6-31G\*\*) for the redistribution reaction to heteroleptic [(BDI)Ca(HMDS)] complexes. Note that the values are given for the formation of 4 moles of [(BDI)Ca(HMDS)].

calculations, in combination with the analysis of the molecular structures, suggest that the backbone modification in the *iPrDip*BDI derivatives contribute to the exergonicity of the reaction, for example by the steric crowding and large strain in [(*iPrDip*BDI)<sub>2</sub>Ca] **7**, which is expected to lower the  $\Delta H$  for the reaction by destabilising the homoleptic species. The strain may further contribute to possible differences in kinetic hurdles between both systems, which were not computationally studied, such as more facile decoordination of one N(Dip) arm at higher temperatures for the bulkier system, e.g. after initial (HMDS)<sub>2</sub>Ca⋯(H)C(C*iPr*(NDip))<sub>2</sub>Ca backbone coordination in an intermediate, as part of an anionic ligand exchange mechanism.<sup>53</sup>

### 3. Conclusions

In conclusion, the synthesis of a heteroleptic calcium amide [(*iPrDip*BDI)Ca(HMDS)] **5** via the manipulation of the Schlenk-type equilibrium to destabilise a normally “kinetically inert” homoleptic species [(*iPrDip*BDI)<sub>2</sub>Ca] **7** has been presented. This ligand redistribution is made possible by the thermodynamic destabilisation of the homoleptic species [(*iPrDip*BDI)<sub>2</sub>Ca] relative to the heteroleptic congener, enabled by careful ligand modification. A one-pot system has been developed harnessing this ligand redistribution, which directly affords the heteroleptic calcium amide complex [(*iPrDip*BDI)Ca(HMDS)] **5** from the reaction of the potassium complex [(*iPrDip*BDI)K] **6**, calcium iodide and Ca(HMDS)<sub>2</sub> in a non-coordinating solvent in good yield. Insights into ligand redistribution strategies like in the presented case are also of interest for the stability and possible regeneration of catalytically competent heteroleptic complexes as part of catalytic cycles and to suppress complex decomposition pathways. Complex **5** has also been used as a precursor species to the new donor solvent-free calcium hydride complex [(*iPrDip*BDI)CaH]<sub>2</sub> **9**, which has been found to be substantially more thermally stable to Schlenk-type redistribution than [(*MeDip*BDI)CaH]<sub>2</sub>. Further investigation of

the reactivity of compound **9** is ongoing, with the application of this thermally promoted Schlenk-type redistribution towards heavier alkaline earth metals also subject to investigation.

### Author contributions

MdVT: investigation, formal analysis, computational analysis, writing – original draft. AMZS, DBC: X-ray crystallography, validation, data curation. AS: writing, visualisation, supervision, funding acquisition, conceptualisation.

### Conflicts of interest

The authors have no conflicts to declare.

### Data availability

Selected data that support the findings of this study are available in the Supporting Information of this article. CCDC 2514137-2514141 contains the supplementary crystallographic data for this paper. These data can be obtained free of charge from The Cambridge Crystallographic Data Centre via [www.ccdc.cam.ac.uk/structures](http://www.ccdc.cam.ac.uk/structures). The research data (NMR spectroscopy, DFT computational studies) supporting this publication can be accessed at <https://doi.org/10.17630/e377abd1-b331-464b-b555-1214950834ea>.

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### Notes and references

- The reaction is reported to require 4 days for full conversion when heating to 70 °C in benzene; in our hands, we found that effectively quantitative conversion to [(*MeDip*BDI)Ca(HMDS)] **3** was achieved after heating the mixture for 16 hours at 80 °C in toluene.
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## Data availability

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Selected data that support the findings of this study are available in the Supporting Information of this article. CCDC 2514137-2514141 contains the supplementary crystallographic data for this paper. These data can be obtained free of charge from The Cambridge Crystallographic Data Centre via [www.ccdc.cam.ac.uk/structures](http://www.ccdc.cam.ac.uk/structures). The research data (NMR spectroscopy, DFT computational studies) supporting this publication can be accessed at <https://doi.org/10.17630/e377abd1-b331-464b-b555-1214950834ea>.

