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Investigating trends in actinide covalency and magnetism with $^{35/37}\text{Cl}$ SSNMR spectroscopy and first-principles calculations

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Abstract

Solid-state NMR (SSNMR) spectroscopy is a powerful technique for studying actinide chemistry but has been significantly limited due to the complex paramagnetism, and radiological hazards presented by these materials. Lanthanide and actinide salts often feature magnetic ordering and can be paramagnetic, ferromagnetic, or antiferromagnetic depending on temperature and electronic structure. Paramagnetic interactions can manifest in SSNMR both as secular spectral shifts and/or couplings as well as contributions from non-secular relaxation. Both effects can be directly measured with NMR and used to extrapolate rich chemical information such as coordination environments, bonding characteristics, local molecular dynamics, and correlation times. Typically, these studies are carried out on high- γ and highly abundant NMR-active isotopes (*e.g.*, ^1H , $^6/7\text{Li}$, ^{19}F , ^{23}Na , etc.) or on enriched rare isotopes (*e.g.*, ^2H and ^{17}O), which can be expensive. Herein, we present a facile methodology to measure the $^{35/37}\text{Cl}$ electric-field gradient (EFG) and paramagnetic shift anisotropy (SA) tensor components using static wideline SSNMR measurements of LaCl_3 , NdCl_3 , UCl_3 , and UCl_4 . The static powder spectra were measured with both ^{35}Cl and ^{37}Cl SSNMR to increase the fidelity of the extracted tensor parameters. Variable temperature NMR of a select case confirms the Curie-Weiss paramagnetism. Relaxation measurements of both nuclei further corroborate observations owing to the paramagnetic relaxation enhancement and reveal simultaneous quadrupolar relaxation mechanisms. Density functional theory (DFT) calculations using Hubbard U corrections to the uranium valence orbitals show excellent agreement with experimental EFG tensor parameters and help describe the bonding characteristics in these lanthanide and actinide systems.



1. Introduction

Actinide salts are a crucial component of molten-salt reactors (MSRs) where chloride salts are used as fuel sources for fissile isotopes (*e.g.*, ^{235}U). The preparation of actinide chloride salts for MSRs currently lacks reliable and quantitative characterization techniques.^{1–3} This limitation is largely due to (i) the hazards presented by radioactive actinide materials and (ii) the complicated physics of *d* and *f* electron orbitals involved in bonding that are reflected in available spectroscopies. The larger radial extent of *d* and *f* orbitals in actinides compared to the lanthanides has also been suggested to allow greater electron sharing with the valence electronic orbitals of ligands. Extensive studies combining X-ray based spectroscopies and *ab initio* modeling have explored these interactions, but direct modeling of these spectra is complex and open to interpretation.^{4–8} NMR spectroscopy is one of the best available techniques for routine characterization and can provide information on local and longer-range structure and provides a direct measurement of the local electron density calculated by density functional theory (DFT). The measured NMR parameters are highly correlated to coordination environments and local bonding phenomena and can provide an insight to the complex valence systems of *d* and *f* orbitals in lanthanides and actinides.

Paramagnetic solid-state NMR (SSNMR) spectroscopy is a growing field that can provide rich information about coordination environments, bonding characteristics, and local molecular dynamics. Paramagnetism manifests in NMR due to the coupling of unpaired electron spins to nuclear spins, referred to as the hyperfine coupling interaction as described by the hyperfine tensor, \mathbf{A} , with the characteristic coupling constant, A . The interaction consists of an isotropic Fermi-contact component, \mathbf{A}_{iso} , which causes isotropic shifts in the resonance frequency of NMR spins, and a pseudo-contact anisotropic electron-nuclear spin dipolar



coupling component, A_D , which causes inhomogeneous broadening in these resonances. Both of these components can experience further inhomogeneous broadening when bulk magnetic susceptibility (BMS) effects are present.⁹ The shifts and inhomogeneous broadening in the resulting NMR spectra can be quantified by the hyperfine tensor parameters, which are often summarized with shift anisotropy (SA) tensor parameters that reflect the sum of diamagnetic and paramagnetic contributions.^{9–15} Pell *et al.* have provided a comprehensive overview of solution- and solid-state paramagnetic NMR spectra with many case studies.¹⁶ Most of these studies are carried out with favorable NMR handles, *i.e.*, spin-1/2 and highly abundant nuclei, where studies with quadrupolar (spin > 1/2) NMR handles are less common by comparison. Furthermore, there are many SSNMR studies of lanthanide materials but only limited reports for actinide materials.^{17,18,27–30,19–26} Part of the limitation in the actinide studies is the hazard presented by radioactivity.³¹

Over the past 20 years there has been an increasing interest in the use of NMR-active quadrupolar halogens as SSNMR handles (*i.e.*, $^{35/37}\text{Cl}$, $^{79/81}\text{Br}$, and ^{127}I). They are extremely sensitive to local structural changes and dynamics, as reflected in their measured electric-field gradient (EFG) tensor parameters. Of these handles, ^{35}Cl is by far the most accessible for SSNMR studies. Although it has a low gyromagnetic ratio, it has a high natural abundance (75.78 %) and a significantly smaller quadrupole moment, Q , than $^{79/81}\text{Br}$ or ^{127}I , causing the powder spectra to span reasonable breadths at high magnetic field strengths. $^{35/37}\text{Cl}$ SSNMR studies have been carried out on a plethora of organic and inorganic materials as highlighted in several review articles by Bryce and coworkers.^{32–35} Methodologies for studying $^{35/37}\text{Cl}$ and similar NMR handles are continuing to improve with advances in static wide-line and ultra-wide-line NMR,^{36–44} fast-MAS and indirect-detection techniques,^{45,46} multiple-quantum



approaches under MAS,^{47–49} among others.⁵⁰ There have been previous ³⁵Cl SSNMR studies of single crystal metal chloride salts with reported EFG and hyperfine tensor parameters, where the metals include paramagnetic Ni²⁺, Cu²⁺, Mn²⁺, Co²⁺, and Fe²⁺ (depending on temperature).^{51–55} However, there have been comparatively few reports of ³⁵Cl SSNMR of crystalline powder materials that are paramagnetic,^{56–58} with the exception of numerous perchlorate and chlorate samples.³³ Bauder and Wu recently reported ^{35/37}Cl SSNMR of paramagnetic Co²⁺ chlorides and demonstrated substantial shifts observed in ^{35/37}Cl spectra.⁵⁹ Recently there have also been nuclear quadrupole resonance (NQR) studies of uranium complexes using ^{87/85}Rb and ³⁵Cl as spectroscopic handles, where only the influence of the EFG tensors can be measured in those cases.^{60,61} There have been considerable efforts in the analysis of ²H SSNMR as simultaneously influenced by the first-order quadrupolar interaction (FOQI) and SA primarily using modified echo and Car-Purcell/Meiboom-Gill (CPMG) type pulse sequences.^{62–69} The only comparable approaches of separating second-order quadrupolar interactions (SOQI) and SA interactions in half-integer spin nuclei are with the COASTER pulse sequence from Grandinetti *et al.*, which needs to be performed under MAS with a rotor angle of 70.12°,⁷⁰ and the TOP-STMAS/MQMAS method from Carvalho *et al.*, which also requires MAS.⁷¹

When studying half-integer-spin quadrupolar nuclei such as ^{35/37}Cl ($I = 3/2$), researchers often focus on the acquisition of the central transition (CT) powder pattern, which spans a far narrower frequency breadth compared to the satellite transitions (STs). CT powder patterns can be simultaneously influenced by the SOQI, SA interactions, and/or dipole-dipole interactions; herein, we focus on simultaneous SOQI and shielding interactions. The complicated orientation dependence of the SOQI convolved with the effects of SA results in static powder spectra that are difficult to analyze and deconvolve their individual contributions.⁷² The problem in these



cases is that eight parameters are needed to accurately model the EFG and SA tensor parameters (two for the EFG, three for the SA, and three Euler angles relating the orientations of the two tensors). To accurately determine all eight parameters, multiple NMR spectra are required. Typically, a pattern should be recorded under MAS, in which the manifestation of the SA parameters is averaged, and a static pattern should be measured. Then both should be acquired at a second magnetic field such that the manifestation of the EFG and SA parameters will change. One problem with this strategy is for cases where the inhomogeneous broadening (from EFG, SA, or both) is too large for sufficient averaging of under MAS, where complex spinning sideband manifolds can hamper analysis. Furthermore, access to a second magnetic field is not a ubiquitous option for researchers and in the current work, radioactive samples are measured where limited exposure to multiple pieces of equipment and spectrometers is desirable.

Instead, the use of ^{37}Cl NMR as a compliment to ^{35}Cl measurements can mimic the effects of measuring ^{35}Cl at a second higher-magnetic field. A higher field will cause a decrease in pattern broadening from the SOQI (with a ν_Q^2/B_0 dependence, where ν_Q is the quadrupolar frequency) and a linear increase in broadening from magnetic shielding (linear B_0 dependence). ^{37}Cl will experience a smaller C_Q as compared to ^{35}Cl (as scaled by the ratio of their quadrupolar moments, $Q(^{37}\text{Cl})/Q(^{35}\text{Cl}) = 0.78812$) and the same η_Q resulting in less inhomogeneous broadening from the SOQI and will be influenced by nearly identical SA tensor parameters.^{73–75} Using $^{35/37}\text{Cl}$ together can provide a much simpler experimental strategy for determining multiple tensor parameters since the Larmor frequencies are close together (*i.e.*, a single broadband probe can measure both) and limit exposure for radioactive materials.

Herein, we discuss the use of $^{35/37}\text{Cl}$ SSNMR to study lanthanide and actinide chloride salts. The central strategy for this study is to use static ^{35}Cl SSNMR measurements to establish



the EFG and paramagnetic SA tensor parameters and then confirm them with ^{37}Cl SSNMR measurements. In every case 1D spectra, T_1 's, and T_2^{eff} 's are measured and compared between the compounds studied. We also calculate local molecular orbitals (LMOs) and bond order parameters from periodic DFT calculations to help constrain the origins of the SA and EFG tensors to provide insight into the relative covalency of the metal-chloride bonds.

2. Methods

2.1 Samples

Lanthanum trichloride [Strem Chemicals Inc., LaCl_3] and neodymium trichloride [Strem Chemicals Inc, NdCl_3] were purchased. Depleted uranium trichloride [UCl_3]⁷⁶ and uranium tetrachloride [UCl_4]^{2,3,77} were prepared according to literature procedures. Radioactive UCl_3 and UCl_4 samples were transferred into polychlorotrifluoroethylene (PCTFE) inserts within 4 mm zirconia NMR rotors and non-radioactive samples were packed directly into the rotor sleeves; therefore, measurements of radioactive samples contain less material. Sample purity was confirmed by powder X-ray diffraction (XRD) patterns (see Supporting Information).

2.2 SSNMR Spectroscopy

SSNMR spectra were acquired using a Bruker Avance NEO console and a 9.4 T ($\nu_0(^1\text{H}) = 399.94$ MHz) Bruker wide-bore magnet at resonance frequencies of $\nu_0(^{35}\text{Cl}) = 39.186$ MHz and $\nu_0(^{37}\text{Cl}) = 32.618$ MHz. Two Bruker broadband 4 mm HX magic-angle spinning (MAS) probes were used under static conditions, where one was used for radioactive samples and the other was used for non-radioactive samples to limit the potential for equipment contamination. $^{35/37}\text{Cl}$ RF pulse powers and chemical-shift reference frequencies were calibrated using NaCl (*s*)



with $\delta_{\text{iso}} = 0.0$ ppm. All spectra were recorded with the WURST-CPMG pulse sequence with varied parameters (see Supporting Information).^{36,37,78–80} The temperature reading for variable temperature (VT) experiments was calibrated using the ^{207}Pb SSNMR of PbNO_3 .⁸¹ All other measurements were made under ambient conditions around 20 °C. $^{35/37}\text{Cl}$ effective T_2 's (T_2^{eff}) were measured using the WURST-CPMG time-domain data and T_1 's were measured with a WURST-CPMG/Saturation-Recovery (WCPMG-SR) type pulse sequence using 16 relaxation delays.⁸²

2.3 Processing and Simulations

All data were processed with Python using DESPERATE and NMRGLUE.^{83,84} WURST-CPMG echo trains were coadded into a single echo, multiplied by a Gaussian function, Fourier transformed, then phase corrected with zeroth, first, and second-order phasing⁸⁵ in every case. Subsequent simulations and fitting were conducted in ssNake version 1.3.⁸⁶ Additional SIMPSON simulations are presented in the supporting information.^{87,88}

2.4 Density Functional Theory

Plane wave DFT calculations were carried out using the Vienna Ab initio Simulation Package (VASP) version 5.4.4.^{89–92} The Revised Perdew-Burke-Ernzerhof (PBE) functional⁹³ with both the Hubbard U correction (DFT+ U) was used for calculations of the paramagnetic NdCl_3 , UCl_3 , and UCl_4 materials. The rotationally invariant method of Dudarev *et al.* was used where $U_{\text{eff}} = U - J$,⁹⁴ where J is the on-site exchange interaction and in all cases was set to 0 eV. These materials were calculated in the ferromagnetic state and scaled to the paramagnetic regime using a Curie-Weiss model as described in Kim *et al.*⁹⁵ The optimal U_{eff} value was obtained by



running calculations for different U_{eff} values and determining the best match to the measured EFG tensors (Supporting Information – **Figures S9-12**). Calculations were also performed including spin-orbit coupling but provided qualitatively inaccurate results for NdCl_3 and had minor effects on those for UCl_3 and UCl_4 (Supporting Information – **Table S3**). Thus, all calculations reported in the main text were performed in the absence of spin-orbit coupling for consistency. All calculations were performed on the un-relaxed XRD structures drawn from the literature. A plane-wave kinetic energy cutoff of 500 eV and convergence energy difference of 10^{-8} eV were used for EFG and chemical shift calculations. The Local Orbital Basis Suite Towards Electronic-Structure Reconstruction (LOBSTER) software package was used to perform crystal orbital Hamiltonian population (COHP), crystal bond order index (COBI), and reconstructed local molecular orbital (RLMO) analyses.^{96–98}

3. Results and Discussion

3.1 Overview

In the following sections, ^{35}Cl and ^{37}Cl SSNMR measurements are conducted on LaCl_3 , NdCl_3 , UCl_3 , and UCl_4 , which all feature a single chlorine site. Diamagnetic LaCl_3 is used as a control, since it is expected to exhibit CT powder spectra only influenced by the SOQI, which will serve to baseline the EFG tensor parameters for structurally similar analogues. All subsequent samples are paramagnetic at room temperature and are expected to be influenced by paramagnetic SA interactions. ^{35}Cl spectra are used to establish both the EFG and SA parameters using fits in ssNake and ^{37}Cl spectra are used to confirm the fitted parameters. Variable temperature (VT) ^{35}Cl NMR is used to verify the paramagnetic SA parameters in a select case for NdCl_3 . In every case the $^{35/37}\text{Cl}$ T_1 's and T_2^{eff} 's are measured and compared. DFT calculations



help constrain the fitted EFG and SA tensor parameters and provide insight to the type of bonding in these lanthanide and actinide chloride compounds.

3.2 $^{35/37}\text{Cl}$ SSNMR

^{35}Cl SSNMR was conducted on all samples outlined above (**Figure 1**). The first compound considered is LaCl_3 (**Figure 1a**), which features a single chlorine site that is in a 9-coordinate geometry around La^{3+} .⁹⁹ The static powder pattern is fit in ssNake where the EFG and SA tensor parameters are extracted. The fitted parameters for this and all compounds are summarized in **Table 1**, where LaCl_3 has $C_Q = 7.91$ MHz, $\eta_Q = 0.49$, $\delta_{\text{iso}} = 302.3$ ppm, and all other SA parameters are negligible. There is no discernable influence of SA on the static powder pattern within the precision of our measurements. The δ_{iso} reported is well within the range of diamagnetic shielding that is often reported for ^{35}Cl SSNMR of metal complexes and ionic compounds.^{32–34} These tensor parameters and the visual appearance of the spectrum and its discontinuities are highly commensurate with a typical static SOQI CT powder pattern, which is consistent with the material being diamagnetic.

NdCl_3 was then investigated, which is structurally similar to LaCl_3 since both compounds occupy a $\text{P6}_3/\text{m}$ space group, are both 9-coordinate to the metal center, and have very similar M-Cl bond distances (2.95 and 2.97 Å for LaCl_3 ; 2.90 and 2.95 Å for NdCl_3).⁹⁹ The main difference is the unpaired $4f$ electrons of Nd^{3+} . The resulting ^{35}Cl pattern (**Figure 1b**) is significantly different from that of LaCl_3 where similar EFG parameters are found with $C_Q = 8.99$ MHz, $\eta_Q = 0.48$, but significant SA parameters are necessary to fit the spectrum with $\delta_{\text{iso}} = 785.5$ ppm, $\Omega = 2004$ ppm, and $\kappa = 0.52$. For comparison, the pattern was fitted without SA parameters, which causes a significant mismatch in the central discontinuity and the broadening of the “horn”



features in the powder pattern (**Figure S5**). This δ_{iso} and Ω are amongst the largest reported for ^{35}Cl NMR and are unlikely to arise from diamagnetic shielding. Since NdCl_3 is paramagnetic, it is very likely that the SA tensor parameters are primarily from the manifestation of paramagnetic shielding, especially as compared to the diamagnetic LaCl_3 , which is almost identical in structure.

UCl_3 is also structurally similar to LaCl_3 and NdCl_3 (*i.e.*, they have the same space group, coordination environment, and similar bond distances)¹⁰⁰ but like NdCl_3 , U^{3+} features unpaired f electrons ($5f$ in this case) and is paramagnetic at room temperature. The ^{35}Cl pattern of UCl_3 (**Figure 1c**) is visually much more similar to that of NdCl_3 rather than LaCl_3 and has similar EFG tensor parameters to both with $C_Q = 8.68$ MHz and $\eta_Q = 0.52$. The SA parameters are significantly larger than both LaCl_3 and NdCl_3 with $\delta_{\text{iso}} = 1014$ ppm, $\Omega = 2255$ ppm, and $\kappa = 1.00$. UCl_3 has a slightly higher paramagnetic susceptibility value than NdCl_3 , which could be one source for the larger SA tensor components, and it is structurally very similar to NdCl_3 , which is reflected in the EFG tensor parameters.

UCl_4 is structurally unique from the trichloride compounds analyzed above, occupying an $I4_1/amd$ space group and having an 8-coordinate geometry around U^{4+} .¹⁰¹ The ^{35}Cl powder pattern (**Figure 1d**) is substantially broader than the others, owing to both the EFG parameters of $C_Q = 12.0$ MHz and $\eta_Q = 0.35$ and the SA parameters with $\delta_{\text{iso}} = 600$ ppm, $\Omega = 5200$ ppm, and $\kappa = -0.01$. The larger C_Q and substantial Ω cause the large frequency dispersion in this case. The δ_{iso} is less than that of NdCl_3 and UCl_3 , but still much larger than that of LaCl_3 , which matches the relative magnetic susceptibility between all the compounds (*vide infra*).

As mentioned earlier, it is difficult to isolate the EFG and SA parameters simultaneously using only one spectrum at one field and only acquired under static conditions as there are eight



free variables involved in the fit. The ^{35}Cl powder patterns presented herein are far too broad to garner useful information from MAS experiments at moderate speeds and may still be difficult with ultrafast MAS (**Figure S6**). The issue of using MAS would not be alleviated at separate magnetic fields, since the net pattern breadth will increase at lower fields, marginally decrease at some higher fields, and increase at significantly higher fields (**Figure S7**). Static measurements at a second magnetic field would be a useful check and constraint on the parameters; however, our laboratory (like many others) is equipped with only a single SSNMR system, and even with a second system, additional equipment would be exposed to radiological hazards. Therefore, using the ^{37}Cl isotope as a handle for a second reference for the tensor parameters is an ideal alternative for simplicity and safety. Any shift differences experienced by the nuclei should be far less than the precision of our measurements.⁷⁴

The ^{37}Cl NMR spectra were acquired for all the compounds described above (**Figure 1e-h**). In every case, the tensor parameters from the fits of the ^{35}Cl data were used to simulate the ^{37}Cl spectra (except for the CQ, which was scaled accordingly). In each case, the simulated spectra match the experimental well. The main drawback of ^{37}Cl is the slightly lower gyromagnetic ratio and its natural abundance being about one third of ^{35}Cl causing a lower receptivity; therefore, the experimental ^{37}Cl are slightly lower in signal-to-noise ratios and in some cases require slightly more Gaussian broadening. Nonetheless, the main discontinuities between all spectra align well, with the exception of the UCl_4 pattern (**Figure 1h**). This pattern appears to experience a bandpass-type depletion of signal, especially to the low-frequency part of the pattern, making it harder to clearly resolve the lower frequency horn (and its splitting) as well as the low frequency pattern edge, but is still discernable. There is also an anomalous feature in both $^{35/37}\text{Cl}$ patterns of UCl_3 where the high-frequency edge shows a peak-like feature (**Figure**



1c,g). This may be due to T_1 or T_2 anisotropy, which would require much more sensitive measurements to confirm.^{82,102} It is important to note that ^{37}Cl spectra are not used for fitting but serve as a check against the fitted parameters from the ^{35}Cl data, increasing our confidence in both the EFG and SA tensor parameters across all compounds.

Having established that chloride compounds containing elements with non-negligible paramagnetic susceptibilities cause the onset of the effects of SA interactions, it is important to rule out that nuclear magnetic shielding (*i.e.*, diamagnetic shift or chemical shift) is causing these effects. The paramagnetic susceptibility and therefore the SA tensor parameters should exhibit a $1/(T-\Theta)$ temperature dependence, provided the temperature exceeds the Weiss constant, Θ .^{9,16} The VT ^{35}Cl SSNMR experiments were carried out for NdCl_3 since its lack of radioactivity allows us to measure more sample (due to our sample-packing technique).

The ^{35}Cl SSNMR of NdCl_3 was measured between *ca.* $-10\text{ }^\circ\text{C}$ and $+50\text{ }^\circ\text{C}$ (**Figure 2a**). The ^{35}Cl SSNMR spectra were recorded at each temperature using the same acquisition parameters as in **Figure 1b** for the same amount of time (about 1 hour per spectrum). As the temperature increases, a noticeable shift in the high-frequency discontinuities of the ^{35}Cl powder patterns are discernable (*i.e.*, the edge and “horn” discontinuities as marked with dashed lines). However, there is no shift observed in the lower-frequency discontinuities. If the EFG parameters were changing with temperature, which can happen in dynamical systems,^{103–105} then both the high- and low-frequency features of the pattern would be expected to change (*i.e.*, if a larger or smaller effective C_Q was observed due to motional averaging, then the entire breadth of the pattern should grow or shrink, respectively – **Figure S8**). Therefore, it is safe to conclude that only the SA tensor parameters could be changing in this case. The subsequent spectra were fit in ssNake using the established EFG tensor parameters from **Figure 1b** and **Table 1** as a



starting point then the SA tensor parameters alone were fitted at each temperature. A summary of the fitted SA tensor parameters in standard convention are summarized in **Figure 2b**. δ_{iso} is found to shift from 840.2 to 681.4 ppm over the temperature range. The fitted δ_{iso} 's can be fitted with a linear function with $R^2 = 0.989$. Therefore, the δ_{iso} exhibits a $1/(T-\Theta)$ temperature dependence characteristic of the Curie-Weiss paramagnetic shielding. UCl_3 is nearly identical in structure to NdCl_3 and has a slightly higher paramagnetic susceptibility; therefore, it is safe to assume that UCl_3 also experiences the Curie-Weiss paramagnetic shielding.

The relative $^{35/37}\text{Cl}$ isotropic shift is compared across all samples (**Figure 3a**). The shifts are plotted against the effective magnetic moment, μ_{eff} , of each compound. The room temperature μ_{eff} is zero for LaCl_3 and is known for UCl_3 (3.70) and UCl_4 (3.29),¹⁰⁶ but not for NdCl_3 ; therefore a μ_{eff} of $3.52 \mu_{\text{B}}$ was measured at 298 K using a Gouy balance. The shifts generally increase with the μ_{eff} 's where the hyperfine constant, $A \propto \mu_{\text{eff}}$. However, the origin of the shifts in each sample can vary greatly with the changing electronic structure and is considered further in DFT calculations (*vide infra*). This same μ_{eff} trend is not observed in Ω , specifically with the large $\Omega = 5200$ ppm for UCl_4 (**Table 1**), which may suggest a significant dipolar contribution to the hyperfine coupling for this compound.

3.3 Relaxation Measurements

Relaxation measurements were carried out for the $^{35/37}\text{Cl}$ NMR data described above. The T_1 is measured for both nuclei using a saturation-recovery (SR) type pulse sequence with WURST-CPMG signal detection.^{82,107} Inversion recovery experiments are more accurate in part due to the higher variation in range, but SR offers significantly faster measurements.^{108,109} Since every 1D spectrum in this work is measured with the WURST-CPMG pulse sequence, the



effective- T_2 , T_2^{eff} , is obtained by default in every case, which also aids in signal detection by increasing the signal-to-noise of the spectra. It is important to note that multiple aspects of CPMG acquisition affect the measured T_2^{eff} and this is not necessarily a reliable metric when ascribing physical differences between measurements and samples. Many experimental WURST-CPMG or standard (Q)CPMG parameters will impact the measured T_2^{eff} including the echo delay times, RF fields, pulse widths, flip angles, pulse phase, and the accumulated errors in refocusing pulses.^{41,110–114} T_1 's are not susceptible to these aforementioned effects and should be considered a far more reliable relaxation measurement, but since the T_2^{eff} is measured by default herein, it is also analyzed.

The measured $^{35/37}\text{Cl}$ relaxation constants are summarized in **Table 2** and **Figure 3b**. As mentioned above, LaCl_3 is diamagnetic and features no influence from SA and has $T_1(^{35}\text{Cl}) = 7.3 \pm 0.2$ s. This is compared to NdCl_3 with $T_1(^{35}\text{Cl}) = 6.5 \pm 0.3$ ms, which is approximately three-orders of magnitude smaller than LaCl_3 . Based on the 1D NMR spectra across all the compounds, the only contributing NMR interactions are the quadrupolar and SA interactions; therefore, these are also likely the only contributors to the non-secular relaxation. If quadrupolar relaxation was dominant, such a dramatic decrease in $T_1(^{35}\text{Cl})$ would not be expected for the C_Q only differing by about 1 MHz between samples (*i.e.*, for quadrupolar relaxation, $1/T_1 = R_1 \propto C_Q^2$). Therefore, the $T_1(^{35}\text{Cl})$ difference between the two samples is likely characteristic of the paramagnetic relaxation enhancement (PRE), which has $R_1 \propto A^2 \propto \mu_{\text{eff}}^2$. UCl_3 and UCl_4 likewise feature $T_1(^{35}\text{Cl})$ values that are *ca.* 1000-fold smaller than LaCl_3 . The $T_1(^{35}\text{Cl})$ differences between NdCl_3 , UCl_3 , and UCl_4 all correlate well with their μ_{eff} 's (**Figure 3b**) commensurate with PRE. Although not as reliable as the T_1 measurements, the same relaxation trends across these samples are also observed in the $T_2^{\text{eff}}(^{35}\text{Cl})$ values. The difference between the $T_2^{\text{eff}}(^{35}\text{Cl})$ of



LaCl₃ and the other samples is not as stark as the $T_1(^{35}\text{Cl})$'s, likely because of the added secular factors impacting T_2^{eff} .

^{37}Cl relaxation measurements were also made in every case and reflect the same relative differences between the samples. An important aspect to note is the comparison of the T_1 values between ^{35}Cl and ^{37}Cl (**Table 2**). In every case, the $T_1(^{37}\text{Cl}) > T_1(^{35}\text{Cl})$. The dramatic decrease in $T_1(^{37}\text{Cl})$ between LaCl₃ and the other samples, as well as the $T_1(^{37}\text{Cl})$'s of the paramagnetic samples correlating with their magnetic moments (**Figure 3b**), suggests the same PRE effects as ^{35}Cl . Similar to ^{35}Cl , the $T_2^{\text{eff}}(^{37}\text{Cl})$ values match the μ_{eff} trends (**Figure 3b**). As with T_1 measurements it is also observed that $T_2^{\text{eff}}(^{37}\text{Cl}) > T_2^{\text{eff}}(^{35}\text{Cl})$. The differences between ^{37}Cl and ^{35}Cl are the gyromagnetic ratios, γ , and the measured C_Q 's. The former could cause an additional PRE contribution for each spin isotope (*i.e.*, a PRE originating from scaling γ rather than μ_{eff} as discussed above). Whereas the latter could indicate the influence of a quadrupolar relaxation mechanism in these spin systems. This facet is clear in the case of diamagnetic LaCl₃ where the ^{37}Cl relaxation times are longer in the absence of any PRE. Therefore, both the PRE and quadrupolar relaxation mechanisms are likely simultaneously contributing to the total rate of relaxation in these spin systems.

3.4 Density Functional Theory

DFT calculations are first used to unravel the various contributions to the observed $^{35/37}\text{Cl}$ NMR spectra. In all cases, reasonable agreement to the experimentally observed EFG tensor parameters are obtained with the DFT+ U calculations with close agreement for LaCl₃, NdCl₃ and UCl₄ (**Table 1**). Agreement with the experimental tensor values for UCl₃ can be obtained by a + U correction to the $6d$ -orbitals in lieu of the $5f$ -orbitals but at the consequence of the incorrect



electronic structures (see discussion in the Supporting Information). Therefore, $+U$ corrections can only be applied to the $5f$ -orbitals for UCl_3 , which yields a slightly higher deviation from the experimental values.

We also probe the effects of the electron hyperfine field from the paramagnetic metals on the $^{35/37}\text{Cl}$ SA using experimental and calculated results. The paramagnetic shielding tensor, σ_s , is obtained from the calculated hyperfine tensor, \mathbf{A} , and the g tensor, \mathbf{g} , with the following relationship,¹¹⁵

$$\sigma^S = \frac{\mu_B S(S+1)}{3\hbar\gamma_N k_B (T-\Theta)} \mathbf{g} \cdot \mathbf{A} \quad (1)$$

Where, μ_B is the Bohr magneton, S is the electronic spin quantum number, γ_N is the gyromagnetic ratio of the observe nucleus ($^{35/37}\text{Cl}$), k_B is the Boltzmann constant, T is the temperature in K, and Θ is the Curie-Weiss constant. The inclusion of the Weiss constant in equation 1 compensates for the bulk magnetic susceptibility effects. This equation is used to determine the full SA tensor by substituting the associated principal components of the \mathbf{A} and \mathbf{g} tensors following the procedure outlined by Pigliapochi *et al.*¹¹⁵ Here we use the previously reported g -tensors for UCl_3 ,¹¹⁶ the tensor for Nd^{3+} in LaCl_3 ,¹¹⁷ and U^{4+} in CaF_2 .¹¹⁸

The calculated SA results for NdCl_3 , UCl_3 and UCl_4 are given in **Table 1**. With the overwhelming influence of \mathbf{A} and the lack of experimental values, diamagnetic contributions to the SA are not reported, which yields no SA for the pure diamagnetic LaCl_3 case. For NdCl_3 , the DFT calculations substantially overestimate the δ_{iso} using the literature \mathbf{g} values. However, the results for UCl_3 and UCl_4 show relatively higher agreement to the measured $^{35/37}\text{Cl}$ δ_{iso} 's. The calculated \mathbf{A} contribution to the span, Ω , is in good agreement for UCl_3 , and NdCl_3 , and underestimated for UCl_4 . Investigating the relative contributions to the SA (**Table S6**) reveal that



the contact shift contributions substantially contribute to both the isotropic and anisotropic portion of the SA with the sign and magnitude of the g-tensor having the largest impact on the anisotropy.

These calculations also do not explicitly model the effects of thermally accessible excited electronic states that arise from zero-field splitting (ZFS) or crystal-field splitting (CFS) present in *f*-block metals. ZFS is known to impact chemical shifts in solution-state NMR of heavy element containing compounds and has been extensively described theoretically.^{119,120} Promising efforts have been made to incorporate these effects into solid-state NMR paramagnetic chemical shift calculations.¹²¹ However, the contributions of electron-electron repulsion, spin-orbit coupling, and CFS have been shown to contribute nearly equally to the magnetic properties of the actinide elements and can only be disentangled through extensive computational efforts.¹²² Although application of these methods to actinide and lanthanide-containing materials could improve the inaccuracies present in the current study, such developments remain a challenge.

The DFT results can be further used to investigate the differences in the M-Cl bonds in these materials. In all the samples studied, the single Cl site adopts a bridging environment between two metal atoms. In the trichloride compounds, the three bridging M-Cl distances differ by at most 0.3 % for NdCl₃. However, in UCl₄ the two bridging U-Cl bonds can be characterized as a short and a long bond with lengths of 2.64 Å and 2.89 Å, respectively. The differences in these bonding environments are investigated using the results from the LOBSTER analysis.^{96–98} This analysis provides measurements of bond covalency by calculating the overlap in the orbital energies and provides the simple numerical values in the form of the COHP, and a unitless parameter denoted the COBI which is a unitless parameter scaled from 0 to 1 where 0 is a fully



ionic bonding interaction and 1 is completely covalent. This program was also recently updated to calculate local molecular orbitals (LMO) from periodic DFT calculations, which they denote as RLMO.⁹⁸ Both UCl_3 bonds and the longer UCl_4 bonds have similar COBI values of 0.18 and 0.21, respectively which indicate a similar ionicity of these bonds. The COBI value of 0.38 for the shorter UCl_4 bond indicates a slight increase in the covalency of this bond relative to the longer bonds. This increased covalency results from the energy overlap between the Cl $3p$ and U $5f$ orbitals as revealed in the calculated density of states (DOS, **Figure S16**).

Investigation of the RLMO for the Cl atom reveal a potential source of the increased covalency in UCl_4 . Looking at the RLMOs corresponding to one of the Cl π lone pair orbitals for these two compounds (**Figure 4**) indicates that the electron density for this orbital is highly localized close to the Cl center in UCl_3 whereas in UCl_4 the density extends slightly along the longer U-Cl bond. Similar results have been observed in calculations of actinide hexachloride salts and have been invoked to explain relative covalency differences in this series.⁶ Further, in UCl_3 V_{zz} is negative and points directly at the σ lone-pair orbital (**Figure 4b**), which is characteristic of a more ionic bond.¹²³ Contrastingly, V_{zz} for the UCl_4 EFG tensor is positive and lies nearly parallel to the vector between the two U metal centers (**Figure 4d**), which could suggest potential contributions from dative bonding between the U and Cl like that observed for previously in heavy transition-metal chlorides.^{33,34} The increased covalency in UCl_4 likely explains the large Ω , as there will be a larger \mathbf{A}_D contribution when the U f electrons are closer to the ^{35/37}Cl spins. It is unclear why the DFT calculations of \mathbf{A} do not reflect this increase, but it could be linked to absence of spin-orbit paramagnetic effects and ZFS not being adequately captured. However, these results must be interpreted with care since VASP and LOBSTER project onto the electron density predicted onto lm -decomposed orbitals and although spin-orbit



coupling is included in calculations, some differences with respect to the full Dirac equation could change the electron density, including the $5p_{1/2}$ and $6p_{1/2}$ states that overlap significantly with the nucleus.^{124–126} More work incorporating relativistic calculations based on the Dirac equation are likely required to fully understand these bonding interactions. Taken together, the results from the RLMO and DFT analysis suggest the increase in bond covalency for the U-Cl bond in UCl_4 is responsible for both the increase in the $^{35/37}\text{Cl}$ C_Q and the dipolar contributions to the SA.

4. Conclusions

$^{35/37}\text{Cl}$ SSNMR of lanthanide and actinide chloride salts measured under static conditions offers a simple and sensitive measure of the paramagnetic SA effects without needing isotopic enrichment or specialized equipment. Some of the largest $^{35/37}\text{Cl}$ SA parameters are reported for these materials, which is onset from the paramagnetic shielding resulting from hyperfine fields in uranium. The effects of bulk magnetic susceptibility likely introduce less inhomogeneous broadening than what can be resolved experimentally. The interplay of EFG and SA tensors offers a high-level of structural detail in a single experiment and offers an extremely unique spectral fingerprint of the analyte, which will be beneficial for future product characterization and identification. These measurements benchmark $^{35/37}\text{Cl}$ shift ranges for U^{3+} - and U^{4+} - containing chloride compounds, which will further aid in future product characterization. All materials studied herein feature one chlorine site, but compounds or mixtures with multiple sites can still be resolved under static conditions.^{107,127} Static measurements at a single field alleviate safety concerns when measuring radioactive actinide materials. Measurements presented herein are relatively quick and simple to execute (*i.e.*, < 4 hours for the longest ^{35}Cl measurement) and



can be improved upon with static probes and larger coils, which may enable more advanced ^{35}Cl - ^{37}Cl correlation experiments.¹²⁸ ^{37}Cl measurements are approximately three-times slower but are ultimately not necessary to perform depending on the application. The size of the SA contribution hinders the CPMG-type acquisition presented herein, since the PRE shortens the T_2 ; however, later actinides such as Pu and Am should exhibit less paramagnetic susceptibility than U^{3+} , which would make the respective measurements of PuCl_3 and AmCl_3 easier than UCl_3 .

DFT calculations using Hubbard $+U$ corrections can help constrain the EFG and \mathbf{A} tensor parameters and can enable easier independent identification of SA tensor parameters. Furthermore, these calculations provide an insight to the relative contributions of ionic and covalent bonding in the U-Cl systems, where experiments and calculations indicate a higher degree of bond covalency in UCl_4 than UCl_3 . Our strategies for fast and reliable static measurements of actinide chloride salts will pave the way for high-throughput analysis useful to the molten salt reactor community and to researchers pursuing synthetic and fundamental advancements in actinide chemistry.

5. Supplementary Material

See the electronic supplementary information for additional simulations, experiments, and experimental details.

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7. Data Availability

All simulation input files, processing functions and scripts, and NMR datasets are available from the authors by request.

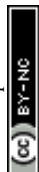


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Table 1: Experimental and Calculated $^{35/37}\text{Cl}$ SSNMR EFG and SA Tensor Parameters^{†,‡,*} and calculated bonding parameters.

Compound		U_{eff} (eV)	$C_Q(^{35}\text{Cl})$ (MHz)	η_Q	δ_{iso} (ppm)	Ω (ppm)	κ (ppm)	α (°)	β (°)	γ (°)	ICOHP (eV)	ICOBI*
LaCl ₃	Exp.		7.91	0.49	302.3	0	0.00	0.0	0.0	0.0		
	DFT	4.0	7.87	0.53	731						-1.13	0.19
NdCl ₃	Exp.		8.99	0.48	758.5	2004	0.52	9.2	4.5	0.0		
	DFT	1.0	8.98	0.49	1738	2549	1.00				-1.17	0.20
UCl ₃	Exp.		8.68	0.52	1014.0	2255	1.00	20.4	2.0	0.0		
	DFT	6.0	9.25	0.58	1071	2038	0.99				-1.10	0.19
UCl ₄	Exp.		12.00	0.35	600.0	5200	-0.01	0.0	8.5	0.0		
	DFT	5.0	-11.99	0.31	-393	3283	-0.43				-1.65	0.29

[†]The principal components of the EFG tensors are ranked $|V_{33}| \geq |V_{22}| \geq |V_{11}|$. The quadrupolar coupling constant and asymmetry parameter are given by $C_Q = eQV_{33}/h$, and $\eta_Q = (V_{11} - V_{22})/V_{33}$, respectively. The sign of C_Q cannot be determined from the experimental $^{35/37}\text{Cl}$ spectra.

[‡]The principal components of the SA tensors are defined such that $\delta_{11} \geq \delta_{22} \geq \delta_{33}$. The isotropic shift, span, and skew are given by $\delta_{\text{iso}} = (\delta_{11} + \delta_{22} + \delta_{33})/3$, $\Omega = \delta_{11} - \delta_{33}$, and $\kappa = 3(\delta_{22} - \delta_{\text{iso}})/\Omega$, respectively. Calculated SA parameters use the full **A** tensor according to Pigliapochi *et al.*¹¹⁵

*The Euler angles α , β , and γ define the relative orientation of the EFG and SA tensors. Euler angles are reported using the $ZX'Z''$ convention.^{129,130}

*The ICOHP and ICOBI values represent bond-weighted averages.

Table 2: $^{35/37}\text{Cl}$ NMR Relaxation Constants

Compound	$T_1(^{35}\text{Cl})$ (ms)	$T_1(^{37}\text{Cl})$ (ms)	$T_2^{\text{eff}}(^{35}\text{Cl})$ (ms) [†]	$T_2^{\text{eff}}(^{37}\text{Cl})$ (ms) [†]
LaCl ₃	$(7.3 \pm 0.2) \cdot 10^3$	$(20 \pm 2) \cdot 10^3$	68 ± 5	650 ± 10
NdCl ₃	6.5 ± 0.3	9.0 ± 0.5	4.4 ± 0.1	6.9 ± 0.3
UCl ₃	2.5 ± 0.4	2.8 ± 0.3	2.3 ± 0.3	2.8 ± 0.2
UCl ₄	10.4 ± 0.9	12 ± 0.1	6.2 ± 0.1	7.8 ± 0.2

[†] T_2^{eff} 's are measured with varying echo delays, RF amplitudes, and WURST sweepwidths (see supporting information).



Figure Captions

Figure 1. ^{35}Cl NMR spectra of (a) LaCl_3 , (b) NdCl_3 , (c) UCl_3 , and (d) UCl_4 and ^{37}Cl NMR of the same compounds, respectively (e – h). Experimental spectra are in black and simulated fitted spectra generated in ssNake are in blue. Spectra are not plotted on the same relative intensity scale.

Figure 2. (a) Experimental variable temperature (VT) ^{35}Cl NMR spectra of NdCl_3 . Vertical dashed lines are drawn to indicate discontinuities in the powder pattern acquired at $-9.88\text{ }^\circ\text{C}$. (b) SA tensor components extracted from fitting each experimental spectrum and plotted in standard convention for visual clarity where $\delta_{\text{iso}} = (\delta_{11} + \delta_{22} + \delta_{33})/3$.

Figure 3. (a) $^{35/37}\text{Cl}$ δ_{iso} 's and (b) $^{35/37}\text{Cl}$ relaxation constants plotted as a function of the effective magnetic moment of each compound studied. $\mu_{\text{eff}} = 0.00, 3.29, 3.52, 3.70$ for LaCl_3 , UCl_4 , NdCl_3 , and UCl_3 respectively.

Figure 4. Graphical representations of the RLMO for (a) the σ lone pair for UCl_3 and (b) the ovaloid representation of the $^{35/37}\text{Cl}$ EFG tensor at the corresponding Cl atom. (c) The RLMO for the π lone pair for UCl_4 and (d) the $^{35/37}\text{Cl}$ EFG tensor. Orange and blue isosurfaces indicate negative and positive charge densities, respectively.



7. Data Availability

All simulation input files, processing functions and scripts, and NMR datasets are available from the authors by request.

