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FEATURE ARTICLE

Engineering Function of Mechanically Interlocked Molecules with Pyrrole-Based Macrocyclic PlatformsRafał A. Grzelczak,^{†a} Daniel Wiendlocha,^{†a} and Bartosz Szyszko,^{*a}Received 00th January 20xx,
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This feature article discusses advances in mechanically interlocked molecules comprising porphyrin or porphyrin-like frameworks, as well as systems in which porphyrinoid structural elements were integrated into interlocked architectures to impart specific functions, such as energy and electron transfer, molecular recognition, catalytic activity, and controlled dynamic operation. The review article traces the field's development from early examples of porphyrin-containing catenanes and rotaxanes to more recent systems in which porphyrinoid stations were deliberately engineered to fulfill defined functional roles. In the final section, the group's contributions to the research area are presented. This includes the progressive development of an iminopyrrole-based self-assembly methodology that enabled the construction of increasingly complex structures – from two-dimensional macrocycles, through capsular assemblies, to metal-stabilized mechanically interlocked molecules. Finally, the synthesis of rotaxanes and catenanes comprising calix[4]pyrins is outlined, demonstrating systems that exhibit a specific type of molecular motion termed fluttering.

Introduction

Over the past few decades, mechanically interlocked molecules (MIMs) have progressed from a synthetic curiosity to a broadly applicable class of functional materials.¹ Their development has enabled the construction of artificial molecular machines,^{2–4} switches,^{5–7} and sensors,^{8,9} as well as structural motifs capable of exerting controlled influence over macroscopic material properties.^{10–12}

Although the fields of pyrrole-based dyes, on the one hand, and mechanically interlocked molecules, on the other, initially developed at different rates and along largely independent trajectories, they began to converge at several key stages.¹³ This convergence became particularly evident in the context of rotaxane research, where structural motifs were sought that were sufficiently large to function as stoppers, while also fulfilling roles beyond purely structural ones. Porphyrin motifs were intended to impart added value to MIMs by introducing versatile coordination properties and attractive photophysical features.^{14–16} Owing, in part, to the seminal contributions of Sauvage, Stoddart, Sanders, Hunter, Anderson, and others, the two areas increasingly informed one another, enabling the transfer of concepts and synthetic strategies.^{17–28}

This article focuses on the points of contact between porphyrinoids and the chemistry of rotaxanes and catenanes. Accordingly, mechanically interlocked systems incorporating classic porphyrin macrocycle are discussed only briefly, particularly in light of the extensive reviews and authoritative accounts by Sauvage and others.^{13–16,29} In contrast, interlocked

architectures built from porphyrinoids^{30,31} – macrocyclic frameworks related to, but distinct from, porphyrins, as well as their derived structural elements – have not yet been systematically reviewed.

This topic has recently become an important line of investigation in our research group at the University of Wrocław. Our interest is motivated by the unusual opportunities that arise from combining pyrrolic macrocycles exhibiting remarkable aromaticity,^{32,33} conformational flexibility,³⁴ intriguing stereochemistry,^{35,36} peculiar reactivity,^{30,37} stabilization of uncommon oxidation states,^{38–40} chemically robust organometallic motifs,⁴¹ and access to multiple electronic states⁴² – with the defining attributes of mechanically interlocked molecules. Mechanical bonding enables control over molecular motion,^{43,44} enforces proximity between interacting components,⁴⁵ and allows tuning of structure, properties, and reactivity that is impossible with conventional covalent design.⁴⁶

The feature article opens with an overview of the field's historical context and a discussion of the work of research groups that laid the foundations for the area. We then outline what we regard as the most important advances in this field, before presenting selected examples of our recent work on the construction of compounds that may, in a broad sense, be described as mechanically interlocked porphyrinoids (MIPs).

Porphyrin-based MIMs

As noted in the Introduction, the use of porphyrins in the construction of MIMs has already been discussed extensively elsewhere; therefore, this topic will be revisited here only

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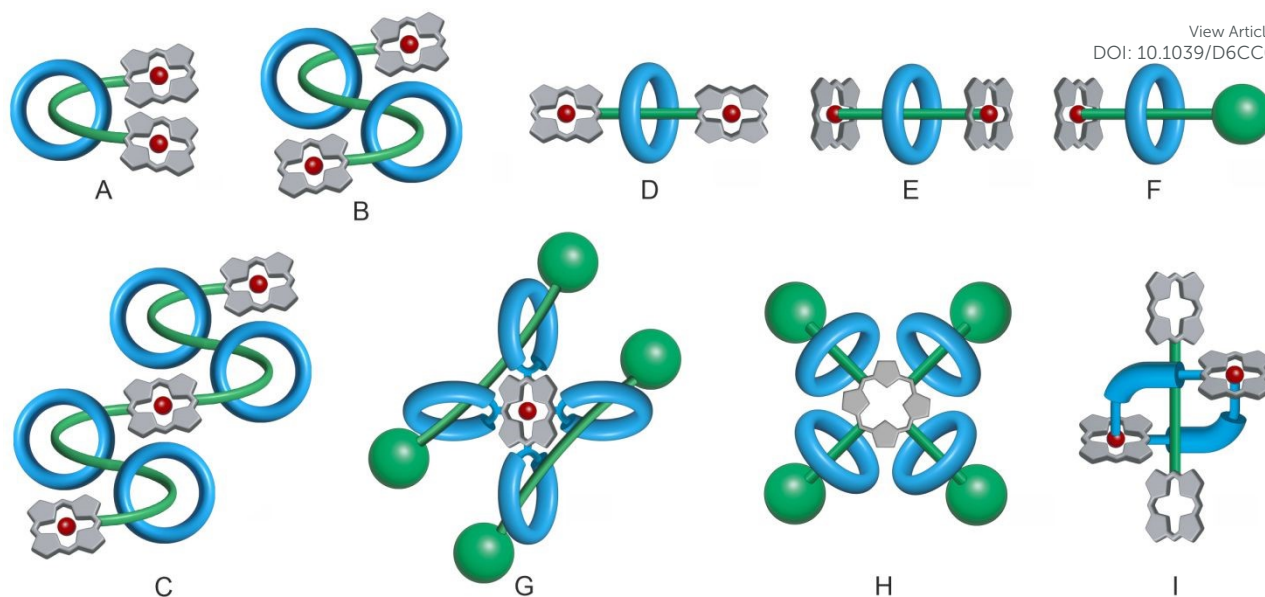


Figure 1. Selected geometric types of rotaxanes and catenanes comprising porphyrin macrocycle(s). Various metals incorporated into the porphyrin cavity were depicted as a red sphere.

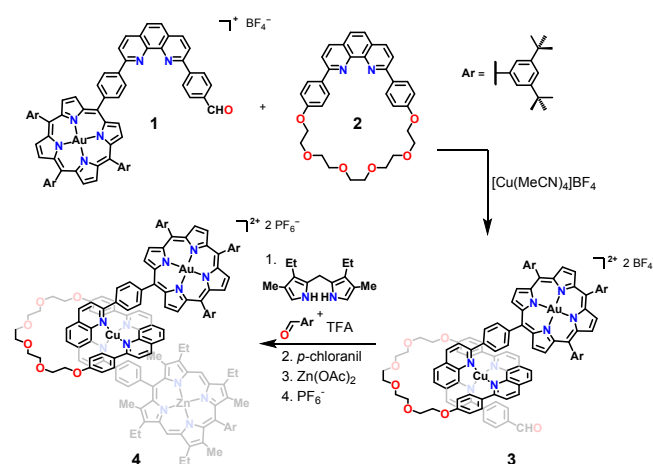
briefly. Nevertheless, any discussion of the connection between pyrrole-based dye chemistry and MIMs would be incomplete without acknowledging the role that porphyrin macrocycles have played – and continue to play – in the supramolecular chemistry of systems stabilized by mechanical bonding.

Porphyrins as structural units

Porphyrins began to be used as structural elements of mechanically interlocked molecules in the 1990s, when their size was recognized as suitable for application as rotaxane stoppers.^{13,17,47,48} Initially, the moieties introduced at the termini of a rotaxane axle served a purely structural role, being intended to prevent dissociation of the macrocycle from the axle, that is, to inhibit the dethreading process. It was later proposed that such entities could exert a far more significant influence, affecting the properties of the interlocked systems through the intrinsic features of porphyrins. From that point onward, the potential of metalloporphyrins was recognized as motifs capable not only of functioning as stoppers but also of introducing intricate features into multicomponent systems.

Using their copper(I)-based passive template approach, the Sauvage group constructed [2]rotaxanes featuring a bent axle bearing two metalloporphyrins at its termini (Figure 1, type A).¹⁷ A central 2,9-diphenyl-1,10-phenanthroline recognition site was incorporated into axle **1**, enabling the formation of a coordination compound upon reaction with a crown ether **2** functionalized with a complementary phenanthroline motif (Scheme 1). The synthesis required the preformation of a copper(I)-stabilized prerotaxane **3** bearing a benzaldehyde functionality at one end of the thread, followed by its condensation with dipyrromethane and 3,5-di-*tert*-butylbenzaldehyde in the presence of TFA. Subsequent oxidation of the initially formed porphyrinogen with chloranil, followed by metalation, afforded the conjugated porphyrin stopper in **4**. The extension of this methodology afforded higher-order systems, including [3]- **5** and [5]rotaxanes **6**

(Figure 1, types B and C, Scheme 2).^{18,49} Removal of the Cu(I) template from the [3]rotaxane by complexation with cyanides afforded a molecule featuring a thread terminated with metalloporphyrins.⁵⁰ The Stoddart group, on the other hand, designed [2]rotaxane **7** composed of the tetracationic blue-box cyclophane assembled around a long polyether chain and terminated with zinc(II) porphyrins connected through their *meso*-aryl positions (Figure 1, type D).¹⁹ Incorporation of two bis(aryloxo) stations within the thread enabled observation of a shuttling process, characterized by a rate constant of 25 s⁻¹ at -10°C and a free-energy barrier of 13.6 kcal mol⁻¹ (Scheme 3

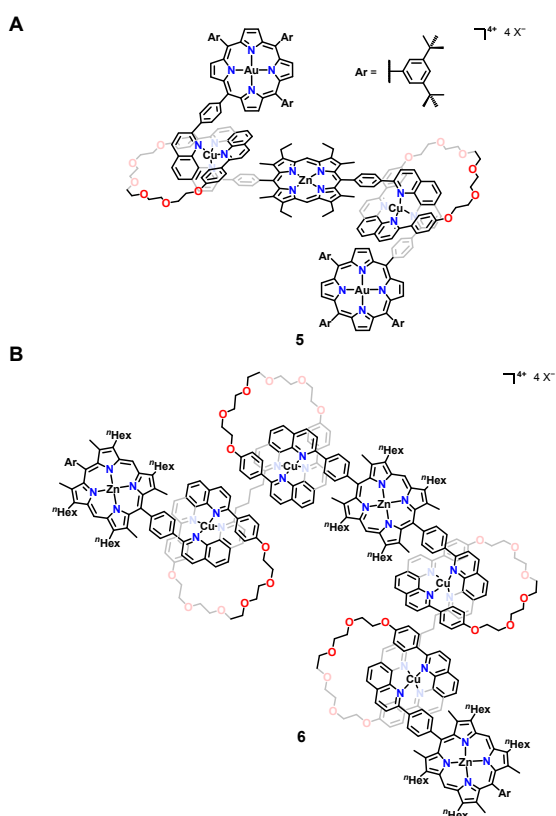


Scheme 1. The synthesis of [2]rotaxane **4**.¹⁷

The affinity of metalloporphyrins toward *N*-donor ligands has been exploited in the construction of rotaxanes in which a *d*-block metal coordination compound of a tetrapyrrole acted as a stopper, coordinatively attached to the termini of a rotaxane axle bearing ligand units such as pyridine (Figure 1, types E and F). The first example of such a system was reported by Branda and coworkers, who designed a self-

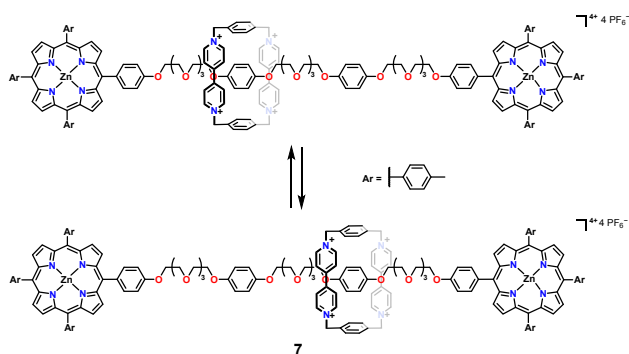


assembling [2]rotaxane composed of dibenzo[24]crown-8 and a



Scheme 2. A) [3]Rotaxane **5** and B) [5]rotaxane **6**.^{18,50}

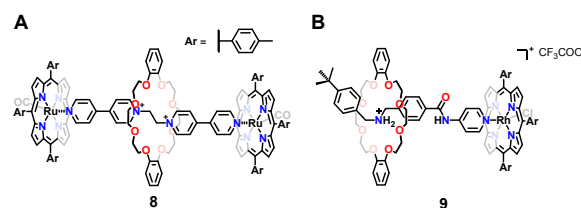
1,2- bis(4,4'-dipyridinium)ethane dicationic thread terminated with 4-pyridyl groups. Upon introduction of carbonyl (tetraphenylporphyrinato)ruthenium(II), possessing one vacant axial coordination site, into a solution of the pseudorotaxane, end-capping occurred at the termini of the axle, transforming the pseudorotaxane species into a [2]rotaxane **8** through the formation of hexacoordinate Ru(II) complexes. Addition of a competitive Lewis base, such as pyridine, led to dissociation of the resulting [2]rotaxane (Figure 1, type E, Scheme 4A).⁵¹



Scheme 3. [2]Rotaxane-based molecular shuttle **7**.¹⁹

Axial coordination to metalloporphyrins containing Zn(II), Ru(II), and Rh(III) was also exploited to introduce stopper groups onto the axle of a [2]rotaxane self-assembled from a macrocycle incorporating electron-rich 1,5-alkoxynaphthalene motifs and a

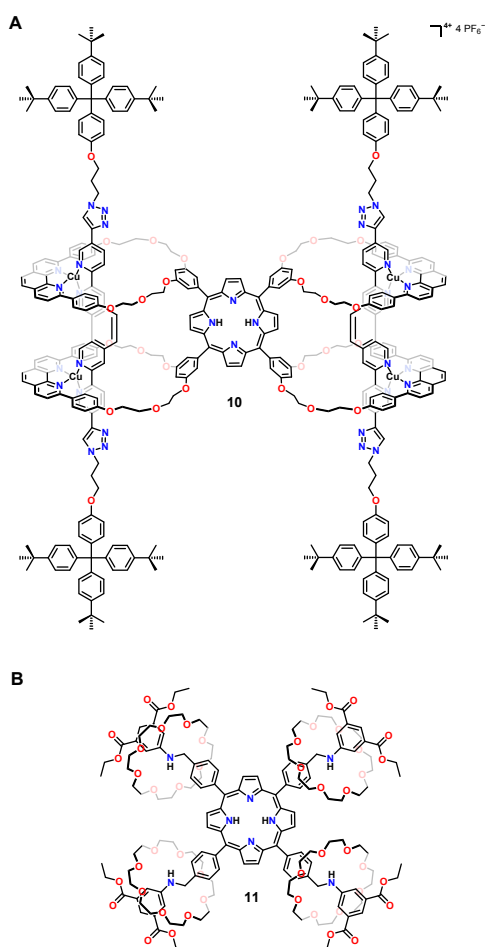
linear thread featuring a central naphthodiimide moiety.²¹ Shimizu systematically examined how variations in the porphyrin macrocycle substitution pattern and the dimensions of the thread influence conformations of [2]rotaxanes stoppered with rhodium(III) porphyrins. One such example is rotaxane **9**, which in contrary to **8**, is terminated with a rhodium(III) porphyrin only on one side (Figure 1, type F, Scheme 4B).⁵²



Scheme 4. [2]Rotaxanes **8** and **9** incorporating metalloporphyrin stoppers.^{51,52}

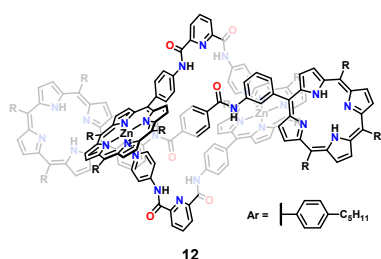
Meso substituents oriented at approximately 90° around the rigid, planar porphyrin core enabled this macrocycle to serve as a platform for constructing higher-order rotaxanes. Sauvage and Sour exploited this design by integrating porphyrin *meso*-aryl substituents into a crown ether framework containing a phenanthroline unit, yielding a Zn(II) porphyrin that enabled the synthesis of a [3]rotaxane **10** via threading of a long axle through macrocycle pairs at the 5,10 and 15,20 positions, followed by stoppering (Figure 1, type G, Scheme 5A).⁵³ By contrast, the Loeb group introduced linear secondary amine groups at the *meso* positions of a porphyrin ring. Using a simultaneous threading-followed-by-stoppering strategy, they obtained a [5]rotaxane **11** based on the tetrapyrrolic scaffold (Figure 1, type H, Scheme 5B).⁵⁴





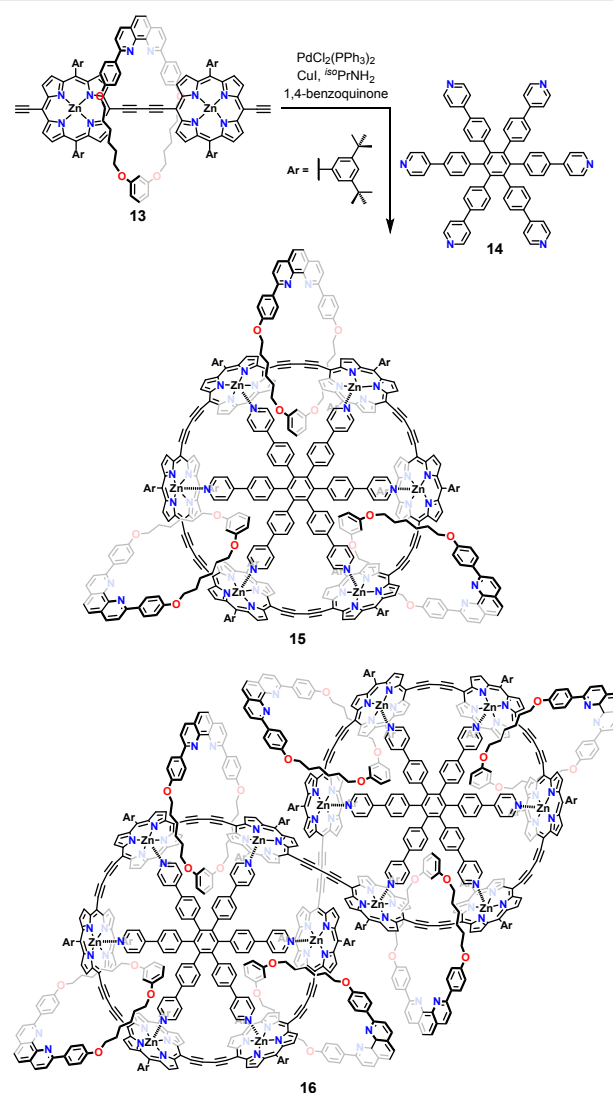
Scheme 5. A) [3]Rotaxane **10** and B) [5]rotaxane **11**.^{53,54}

Hunter designed a zinc(II) porphyrin bearing a carefully engineered *meso* substituent incorporating pyridine termini and a pyridine-2,6-dicarboxamide motif. In chloroform, this species existed in dynamic equilibrium with a self-assembled dimer, in which the pyridine unit of one porphyrin acted as an axial ligand to the Zn(II) center of the other. Owing to cooperative binding, the dimer displayed a very high association constant ($K_d = 2 \times 10^8 \text{ M}^{-1}$). The presence of hydrogen-bond donors and acceptors within the *meso*-pyridinedicarboxamide unit further enabled this dimer to form [2]rotaxanes **12** with a range of terephthaloyl derivatives, in which the porphyrins constituted part of the self-assembled macrocyclic component and served as stoppers at the termini of the thread (Figure 1, type I, Scheme 6).²⁶



Scheme 6. [2]Rotaxane **12**.²⁶

The Anderson group pushed the concept of using porphyrins as components of large macrocyclic architectures to the extreme by designing a [2]rotaxane **13** with a butadiyne axle terminated by two Zn(II) porphyrin stoppers, further functionalized with alkyne groups (Scheme 7).²⁷ The resulting building block was subjected to Pd-catalyzed Glaser coupling in the presence of a hexapyridyl template **14**, affording remarkable [4]- **15** and [7]catenane **16** architectures featuring large ring components composed of cyclic porphyrin hexamers and dodecamers in 62% and 6% yield, respectively. Later, the same group demonstrated several other impressive systems in which the porphyrin was incorporated into the axle or ring component of the interlocked molecule.^{55,56}



Scheme 7. Synthesis of [n]catenanes **15** and **16**.²⁷

Photosynthesis-mimicking systems

Molecular systems incorporating metalloporphyrins were recognized early on as valuable models that mimic the action of photosystems, which convert light into chemical energy.^{13–16} In particular, bis-porphyrin compounds containing a zinc(II) porphyrin unit at one terminus and a gold(III) porphyrin unit at the other were shown to exhibit long-range photoinduced



charge separation, with the zinc(II) porphyrin acting as the excited-state electron donor and the gold(III) porphyrin serving as the electron acceptor.⁵⁷ It was later demonstrated that electron and energy transfer processes can be tuned by incorporating the active species into rotaxane or catenane architectures, which allow for precise spatial separation of the interacting components without close contact through covalent bonds. This structural arrangement suppressed rapid charge recombination, which would otherwise render the processes fast and inefficient.¹⁷ The extensive work of the Sauvage group demonstrated that the activity of metalloporphyrin-stoppered rotaxanes can be finely tuned by modifying the MIM architecture, enabling the construction of multicomponent systems with highly desirable photoinduced electron transfer characteristics.^{58–61} Flamigni, Serroni, and co-workers reported a series of [2]catenanes composed of polyether-strapped zinc(II) porphyrins and cyclobis(paraquat-*p*-phenylene), which exhibited not only efficient photoinduced electron transfer but also a pronounced dependence of the electron-transfer rate on the strap size.⁶²

Multiple systems incorporating C₆₀ fullerene as an electron acceptor were designed, in which photoinduced electron transfer from the excited-state porphyrin led to charge-separated states.^{63–66} Takata, Ito, and co-workers reported an interesting [2]rotaxane **17** featuring zinc(II) porphyrin donors installed at the termini of the thread and a cyclic component decorated with a C₆₀ fullerene acceptor (Figure 2). This design enabled the observation of intramolecular electron transfer between the donor metalloporphyrins and the [60]fullerene in benzonitrile.⁶⁷ Schuster and colleagues reported a [2]rotaxane architecture comprising metalloporphyrin stoppers that acted as electron donors and a macrocycle appended with C₆₀, assembled using the Sauvage bis(phenanthroline)copper(I) motif.⁶⁸ It was demonstrated that the intramolecular electronic processes observed in this system required the involvement of the [Cu(phen)₂]⁺ complex in the electron transfer events. Later, the role of the same motif was documented for [2]catenanes in which metalloporphyrin and fullerene were spatially significantly separated.⁶⁹

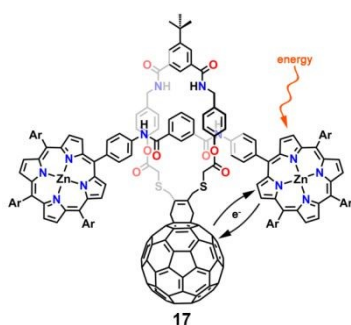
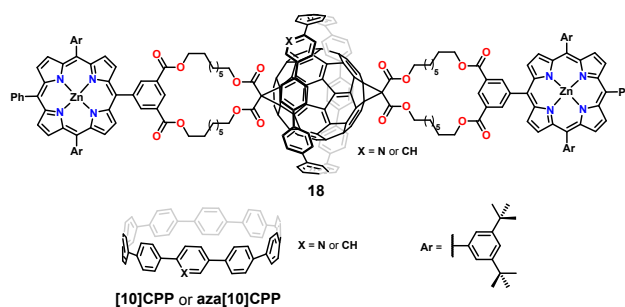


Figure 2. Photoinduced electron transfer in C₆₀-appended [2]rotaxane **17**.⁶⁷

Von Delius and Guldi constructed [2]rotaxane **18** featuring zinc(II) porphyrin stoppers and a C₆₀ fullerene acceptor positioned at the central part of the thread (Scheme 8).⁷⁰ While electron transfer between the donor and acceptor was observed for the free thread, the introduction of a

[10]cycloparaphenylene ring (CPP or its aza-analogue aza[10]CPP) encircling the C₆₀ unit effectively shielded the acceptor. This structural modification suppressed electron transfer and redirected the photophysical pathway toward triplet-state energy transfer between the zinc(II) porphyrin and the fullerene. The Anderson group, on the other hand, described [3]- and [5]-rotaxanes in which multiple energy transfer pathways involving a phenanthroline-based macrocycle, a polyyne thread, and porphyrin stoppers were operative.²⁸

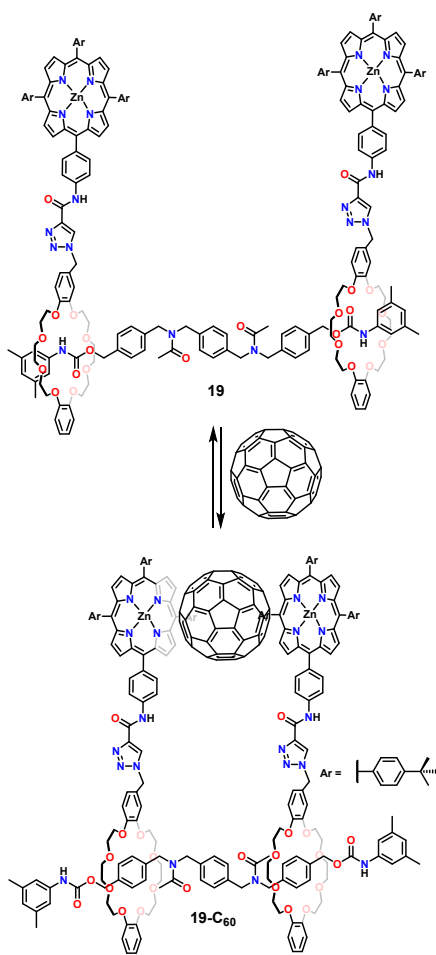


Scheme 8. The [10]CPP-based [2]rotaxane **18** comprising a fullerene moiety.⁷⁰

Interlocked receptors and sensors

Porphyrins have several key features – they are large, aromatic, and planar molecules that, upon metalation of their cavity, can bind axially to N- or O-donor ligands.^{71,72} Combined with the conformational flexibility of rotaxanes, these properties enable the formation of adaptable receptors capable of interacting with a wide range of molecular and ionic guests.⁷³ It was this complementarity that enabled Morin and co-workers to design rotaxanes functioning as tweezer-like receptors for fullerenes (Scheme 9).⁷⁴ The attachment of porphyrin units to two dibenzocrown ether macrocycles led to the formation of a [3]rotaxane **19** in which the distance between the aromatic planes could be modulated. As a result, the porphyrin panels – well known for their efficient interactions with fullerenes – were able to move apart and effectively encapsulate C₆₀, C₇₀, and C₈₄.





Scheme 9. Binding of C₆₀ by the tweezer-like [3]rotaxane receptor **19**.⁷⁴

The Sauvage group further explored the dynamic behavior of [3]rotaxanes featuring two zinc(II) porphyrin plates that interacted with bispyridyl podands differing in linker flexibility between the coordinating units.⁷⁵ The large accessible amplitude of distance variation between the metalloporphyrin panels (10 – 80 Å), together with the cooperative axial coordination of a single podand molecule to two metal centers, enabled exceptionally high complex stability constants, reaching values of 10⁶–10⁷ M⁻¹.

The Beer group has extensively studied anion binding by porphyrin-based rotaxanes and catenanes. By exploiting anion-templation strategies, they developed a range of interlocked receptors capable of recognizing multiple anionic guests, with the porphyrin macrocycle serving as both the recognition site and a reporter unit.⁷⁶ The group also designed a receptor based on a porphyrin-strapped macrocycle in which iodotriazole functionalities were incorporated into the cyclic component of the rotaxane to enable anion binding *via* halogen bonding.⁷⁷ Axial coordination between a triazole unit on the [2]rotaxane thread and a zinc(II) bound in the porphyrin cavity preorganized the binding pocket and led to a substantial enhancement of anion binding affinities. Notably, quenching of a BODIPY unit installed as a stopper at one terminus of the thread enabled optical detection of the anion recognition event. A similar design principle enabled the construction of a catenane **20** that

acted as a host for anionic guests, in which the anions were stabilized within the cavity through a network of hydrogen bonds (Figure 3).⁷⁸

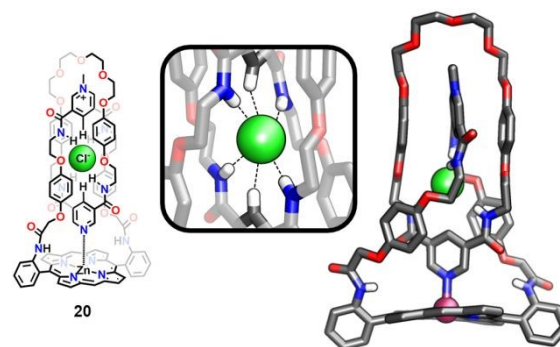
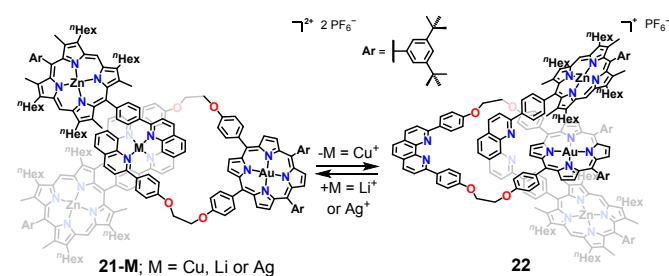


Figure 3. The chloride complex **20** of a porphyrin-incorporating [2]catenane receptor.⁷⁸

Switches and molecular machines

Porphyrin-containing MIMs have been extensively explored as platforms for constructing molecular switches and machines.^{6,19} The Sauvage group demonstrated early on that a [2]rotaxane **21-M** assembled from a thread bearing Zn(II) porphyrin stoppers and a macrocycle incorporating an Au(III) porphyrin undergoes a pronounced intercomponent conformational rearrangement (Scheme 10).⁷⁹ Removal of the copper(I) template induced an approximately 180° rotation of the axle relative to the macrocyclic component, driven by attractive interactions between Zn(II)- and Au(III) porphyrins in **22**. The same group designed a [3]rotaxane in which two phenanthroline-based macrocycles fused with metalloporphyrins were threaded onto a linear component bearing a complementary Cu(I)-coordinating motif.⁸⁰ Upon removal of the template, a shuttling motion was observed, in which the porphyrin plates move closer together or farther apart in response to interactions with guests capable of axial coordination to the Zn(II) porphyrins.



Scheme 10. Demetallation-induced co-conformational rearrangement of **21-M**.⁷⁹

The Beer group designed molecular shuttles based on [2]- and [3]rotaxanes in which a Zn(II) porphyrin unit was positioned at the centre of the axle.⁸¹ The macrocyclic component incorporated a pyridyl ring. In the absence of competing ligands, shuttling of the macrocycle along the thread was observed, with the Zn(II) porphyrin station influencing the dynamics through axial coordination with the pyridyl unit of the ring component.



Upon the addition of an external base, i.e., pyridine, the shuttling was inhibited.

The coordination properties of metalloporphyrins were also exploited by Terao and co-workers, who designed a remarkable [3]catenane **23** comprising a large, rigid rectangular macrocycle incorporating Ru(II) porphyrins and two smaller rings, each containing two Zn(II) porphyrin units (Figure 4).⁸² Upon introduction of carefully selected ligands such as 3,5-dibromopyridine, DABCO, or bisbipyridyls, multistate switching could be reversibly induced.

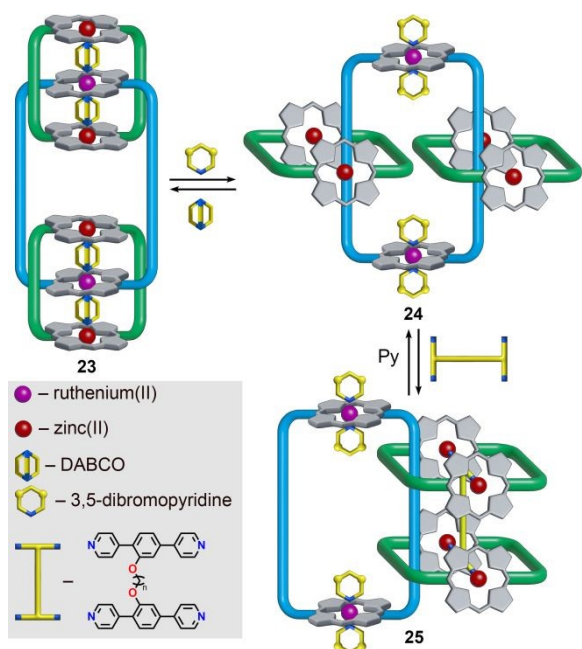


Figure 4. Coordination-induced multistate switching in [3]catenane **23**.⁸²

Catalysts and nanoreactors

In 2003, Nolte and co-workers demonstrated the remarkable potential of MIMs incorporating catalytically active metalloporphyrins.⁸³ A precisely tailored cavity in a hybrid glycoluril–Mn(III) porphyrin system enabled the threading of a polybutadiene polymer strand, which, upon attachment of stoppers, formed a [2]rotaxane **26-Mn** with the polymer serving as the axle. Upon addition of a bulky *tert*-butylpyridine ligand, capable of axial coordination only from outside the cavity, and in the presence of an oxygen source, the double bonds of the threaded polymer were converted into epoxides **27-Mn** (Figure 5). This transformation yielded polybutadiene epoxide, sparking a debate over the mechanism of catalysis.^{83–87}

The Megiatto group, in contrast, demonstrated remarkable reactivity within nanocavities incorporating cobalt(II) porphyrins, forming rotaxanes *via* an original active-template synthetic approach.⁸⁸ In the presence of 3,5-diphenylpyridine, half-threads of the rotaxane bearing diazo and styrene functional groups underwent a radical carbene-transfer reaction, affording the interlocked product of the olefin cyclopropanation in 95% yield. Later, the same group exploited the reactivity of a nanoreactor incorporating a Ru(II) porphyrin

system to demonstrate N–H bond carbene insertion, affording [2]rotaxanes.⁸⁹

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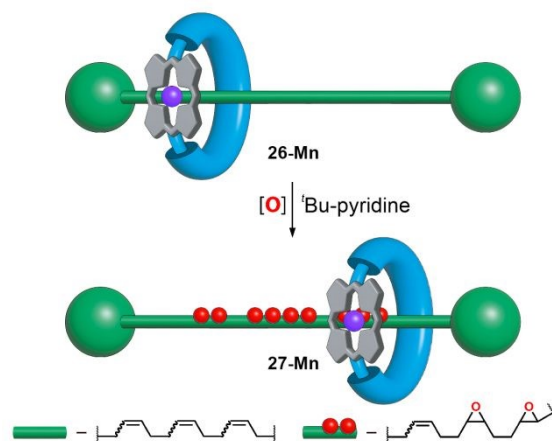


Figure 5. Epoxidation of polybutadiene polymer in a Mn(III) porphyrin-based rotaxane **26-Mn**.⁸³

A particularly interesting concept involving the use of metalloporphyrin-based capsules in tandem active-metal template synthesis of rotaxanes was described by Hayashi, Weiss, and co-workers.⁹⁰ Using a ditopic Zn(II) porphyrin–phenanthroline–strapped system, it was demonstrated that azide coordination to the Zn(II) centre, enforced by the nanoreactor design, facilitated the Huisgen cycloaddition between azide- and alkyne-functionalized half-threads. This process afforded [2]rotaxanes and opened new avenues for exploring dual-metal catalysis in active-template syntheses of mechanically interlocked molecules.

MIMs incorporating porphyrinoids as functional platforms

Although many examples of porphyrins used as building blocks for mechanically interlocked molecules have been reported, the past decade has seen the emergence of MIMs based on porphyrinoids, featuring porphyrin-like macrocycles or structural motifs such as bipyrrrole, bithiophene, and dipyrromethene. This section focuses on interlocked molecules in which porphyrinoids served as platforms for constructing catenanes and rotaxanes.

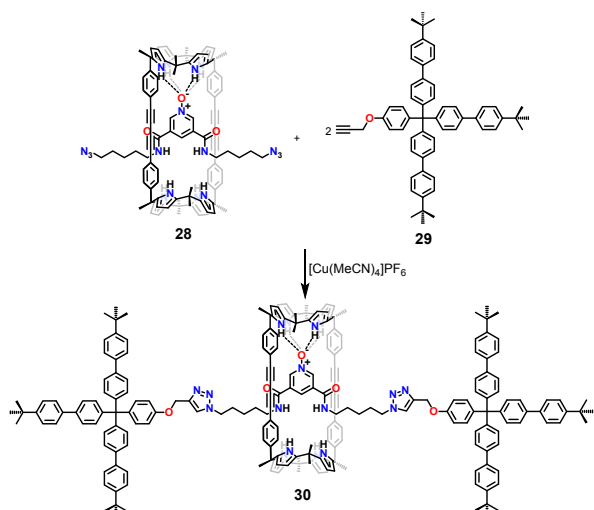
Calix[4]pyrrole-based rotaxanes

The conjoining of four pyrrole rings through tetrahedral bridges yields calix[4]pyrrole, a nonconjugated macrocycle originally synthesized by Baeyer in the nineteenth century and later shown to act as a highly efficient anion receptor.^{86, 87}

In 2012, the Ballester group demonstrated that a cage-like species constructed from two calix[4]pyrrole units connected by rigid bisacetylene linkers is capable of encapsulating a linear zwitterionic guest, bis(amidopyridyl-*N*-oxide), to form a pseudorotaxane **28** (Scheme 11).⁹³ Hydrogen bonding to the pyridinium *N*-oxide oxygen atom oriented the guest such that the amide functionalities were exposed toward the remaining



free space within the cavity. This preorganization facilitated the formation of four-component species upon complexation of tetrabutylammonium cyanate or azide. Installation of bulky stopper groups **29** on the pseudorotaxane **28** using a capping strategy resulted in the formation of a [2]rotaxane **30** in good yield (50%).⁹⁴ This interlocked receptor was shown to bind a range of anionic guests effectively, including chloride and nitrate.^{95,96}



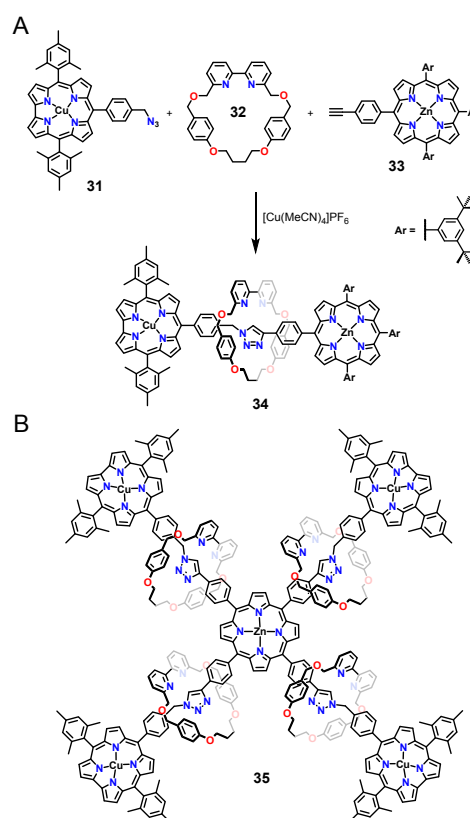
Scheme 11. Synthesis of bis-calix[4]pyrrole [2]rotaxane **30**.⁹⁴

Aydogan and co-workers exploited the anion-binding properties of calix[4]pyrroles to construct alternating polymers composed of linear segments with pillar[5]arenes bearing chains of varying lengths.⁹⁷ The supramolecular polymers were further shown to organize hierarchically into a variety of structures, including spherical aggregates, fibers, microporous films, and three-dimensional materials.

Corrole-embedded interlocked architectures

Corroles are tetrapyrrolic macrocycles that, due to the absence of one methine bridge between the pyrrolic subunits, belong to the so-called contracted porphyrinoids.^{98,99} Ngo, D'Souza, Goldup, and co-workers employed active-template synthesis to prepare a series of mechanically interlocked porphyrin–corrole systems.¹⁰⁰ They demonstrated that the mechanical bond can be exploited to provide steric shielding of porphyrinoid-terminated threads without altering the electronic properties of the porphyrinoids. The copper(I)-catalyzed azide–alkyne cycloaddition reactions between *meso*-functionalized copper(III) corrole **31** and Zn(II) porphyrin **33** bearing an alkyne group at the *meso*-position, in the presence of macrocycle **32**, afforded a [2]rotaxane **34** (Scheme 12A). This methodology was further extended to construct a series of impressive triads and pentads **35** (Scheme 12B). Upon titration of the latter with DABCO, an altered assembly behavior of the interlocked species into dimeric structures was observed. This effect was attributed to the presence of mechanically interlocked substituents at the central Zn(II) porphyrin framework, which acted as a distinctive picket-fence-like architecture.¹⁰⁰ Later, it was demonstrated

that a three-component rotaxation reaction between a metal-free corrole azide, a zinc(II) porphyrin-based alkyne, and a preformed copper(I) complex of a bipyridyl macrocycle, followed by treatment with HBr–PPh₃, afforded corrole–porphyrin conjugates with very good yields.¹⁰¹

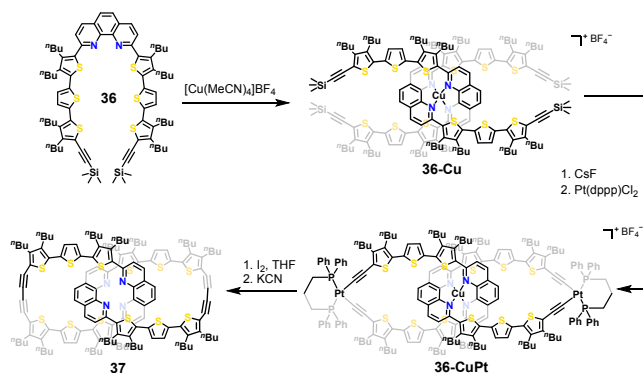


Scheme 12. A) The active template synthesis of corrole-incorporating rotaxane **34**, B) [5]rotaxane **35**.¹⁰⁰

π -Conjugated porphyrinoid catenanes

Cyclic oligothiophenes may be viewed as heteroanalogues of cyclo[*n*]pyrroles – a class of porphyrinoids composed only of pyrrole rings.¹⁰² Given that oligo- and polythiophenes are widely used organic semiconductors with important applications in organic electronics, the properties of interlocked architectures comprising π -conjugated rings are particularly intriguing.^{103,104} It motivated Bäuerle and co-workers to explore the synthesis of catenanes composed of cyclic oligothiophenes and to examine whether two interlocked π -conjugated rings would exhibit mutual electronic communication.¹⁰⁵ The group initially employed a double-templating strategy. Using Cu(I), two oligothiophene-substituted phenanthroline **36** units were arranged perpendicularly in a Sauvage-type geometry **36-Cu** (Scheme 13).¹⁰⁶ Subsequent complexation of the terminal α -ethynyl groups on the thiophene rings to Pt(II) under high-dilution conditions, followed by iodine-mediated reductive elimination, yielded a new C–C bond enabling catenane **37** formation. Although the target compounds were obtained, the synthetic route was rather elaborate and afforded only small amounts of the desired material. The group therefore further refined the methodology and found that significantly improved

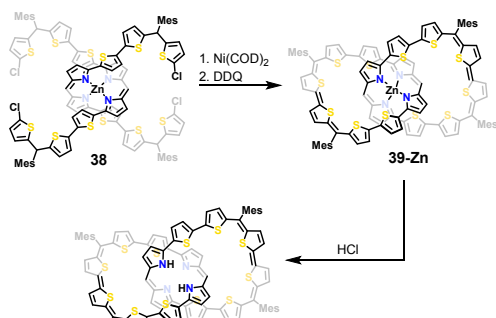




Scheme 13. Synthesis of a conjugated [2]catenane **37**.^{106,107}

yields of conjugated catenanes could be achieved *via* Pt(II)-catalyzed coupling of the terminal ethynyl groups.¹⁰⁷ This development opened access to conjugated catenanes of various sizes. Notably, detailed optical and electrochemical studies indicated that the two macrocycles in the catenane influence one another *via* through-space donor–acceptor interactions.

In 2025, Zeng, Wu, Ni, and co-workers reported the formation of a [2]catenane **39** composed of two intertwined octaphyrinoid macrocycles incorporating both pyrrole and thiophene units.¹⁰⁸ The synthesis began with an α -bromothiophene-substituted dipyrromethene, which, upon coordination to Zn(II), formed a dimeric complex in nearly quantitative yield. Subsequent Suzuki–Miyaura coupling of this intermediate with a dithiophenemethane-derived boronic ester afforded a precursor **38** that was subjected to Yamamoto coupling, yielding the Zn(II)-stabilized catenane **39-Zn** in 24% yield (Scheme 14).



Scheme 14. Synthesis of octaphyrinoid-based [2]catenane **39**.¹⁰⁸

Finally, oxidative dehydrogenation with DDQ followed by demetalation with HCl provided the fully conjugated, metal-free [2]catenane **39**. Single-crystal X-ray diffraction analysis of **39** revealed that the core-modified octaphyrinoid macrocycles were arranged in an almost perpendicular orientation, with all bithiophene subunits adopting *cis* configuration (Figure 6). The aromatic character of both rings was supported by harmonic oscillator model of aromaticity (HOMA) calculations, anisotropy of the induced current density (AICD) analysis, and three-dimensional *iso*-chemical shielding surface (ICSS) maps. Intriguingly, oxidation of **39** to the tetracation induced a pronounced structural distortion, leading to the formation of

two antiaromatic octaphyrinoid units with close inter-ring contacts. This arrangement enabled strong through-space electronic coupling between the macrocycles.

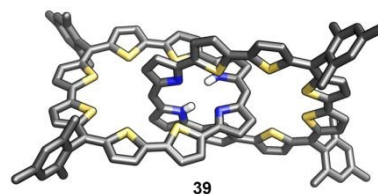
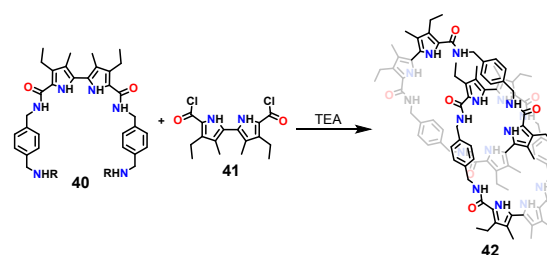


Figure 6. X-ray molecular structure of **39**.¹⁰⁸

MIMs embedding porphyrinoids' structural elements

Several research groups have used not entire porphyrin or porphyrinoid molecules, but rather fragments of these macrocycles, such as pyrrole, bipyrrrole, dipyrromethane, and dipyrin, for the construction of catenanes and rotaxanes. When appropriately functionalized, these units serve as building blocks that act as hydrogen-bond donors/acceptors or as ligands to coordinate metal cations, while simultaneously retaining the ability to form macrocyclic architectures.

In 1998, Sessler and Vögtle reported a spectacular example of a self-assembling [2]catenane comprising bipyrrrole-based macrocycles.¹⁰⁹ The reaction of *p*-xylylenediamine **40** with a bipyrrrole diacyl chloride **41** afforded the interlocked product **42** in 2% yield (Scheme 15). However, the same compound could be obtained in a slightly higher yield (4%) when a sequential synthetic route was employed. The formation of an interlocked architecture was likely facilitated by hydrogen bonds between multiple donor and acceptor sites in the assembling parts of the molecule. The dynamic nature of **42**, involving circumrotatory motion of the individual macrocyclic components, was inferred from the broadening of the ¹H NMR resonance signals. Upon addition of fluoride, the resonances became significantly sharper, consistent with the anion binding within the catenane cavity. Subsequently, the interlocked receptor was shown to effectively interact with other anions as well.

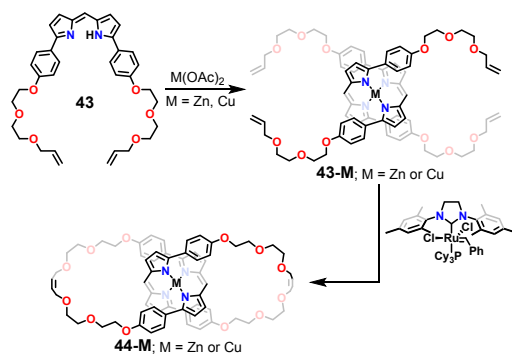


Scheme 15. Synthesis of a bipyrrrole-based catenane **42**.¹⁰⁹

Bozdemir and co-workers exploited the coordination properties of the dipyrin motif to construct [2]catenanes using Sauvage's passive metal-template methodology.¹¹⁰ Accordingly, a dipyrin-embedded macrocyclic precursor **43** bearing terminal alkene functionalities appended to ether side chains was synthesized (Scheme 16). Upon treatment with Zn(II) or Cu(II), ligand dimerization around the tetrahedral metal center



afforded bis(dipyrinato)metal complexes **43-M**, effectively bringing the terminal double bonds into close proximity. Subsequent ring-closing metathesis catalyzed by the Grubbs 2nd-generation catalyst enabled covalent bond formation between the alkene termini, affording **44-M**. Removal of the metal template from **44-M** furnished the metal-free [2]catenane, which could subsequently be transformed into bis-BODIPY-like species.



Scheme 16. Synthesis of **44-M**.¹¹⁰

An interesting strategy for constructing interlocked structures from expanded porphyrin-like macrocycles was reported by Nabeshima.^{111,112} The group described several macrocyclic trimers incorporating dipyrin motifs separated by *meta*- and *para*-substituted phenylene units. Their corresponding tris-BF₂ complexes were capable of incorporating protonated secondary amines to form pseudorotaxanes, e.g. **45**, via non-classical, bifurcated hydrogen bonds, namely BF₂⋯H–N interactions (Figure 7).

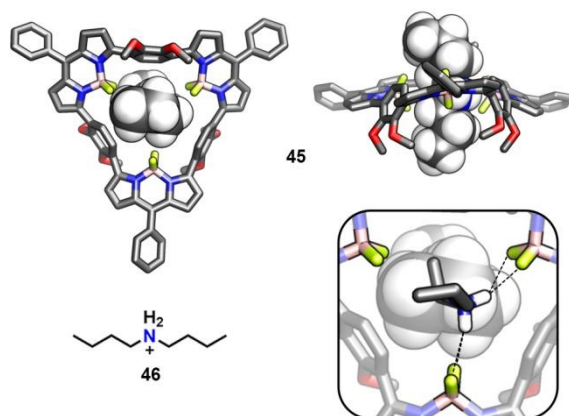


Figure 7. The pseudorotaxane **45** comprising expanded porphyrin-like ring component.^{111,112}

The BODIPY motif has been incorporated into rotaxane architectures on several occasions to exploit its outstanding optical properties, including large molar extinction coefficients and high fluorescence quantum yields. Importantly, these features are complemented by excellent chemical stability and synthetic versatility, which allow fine-tuning of the luminescent properties. Akkaya and co-workers designed a [2]rotaxane in which suitably functionalized BODIPY units were attached both

to the macrocyclic component – dibenzo-24-crown-8 – and installed as stoppers on the axle bearing a secondary amine at its center.¹¹³ Protonation of the amine generated the corresponding ammonium species, triggering spontaneous self-assembly of the [2]rotaxane. The resulting interlocked structure exhibited efficient through-space energy transfer *via* FRET.

Frasconi, Wasielewski, Stoddart, and co-workers reported a bistable [2]rotaxane composed of a tetrathiafulvalene (TTF)-containing thread terminated with a BODIPY stopper and a cyclobis(paraquat-*p*-phenylene) macrocycle.²⁰ In the initial redox state, in the absence of an applied potential, the macrocycle resided on the TTF station, allowing free rotation of the BODIPY unit around the C_{meso}–C_{ipso} bond and resulting in low fluorescence intensity. Upon oxidation of TTF to the dication TTF²⁺, the macrocycle shifted toward the BODIPY stopper. This movement restricted the rotation, resulting in a 3.4-fold increase in fluorescence.

Trolez and co-workers described an intriguing route toward interlocked dipyrromethene derivatives. They designed a macrocycle incorporating a central 2,2'-biphenol unit.^{114,115} When the latter was added to a solution of BODIPY, previously treated with aluminum chloride in anhydrous dichloromethane, substitution of the fluorine atoms by the 2,2'-biphenol moiety took place. This transformation yielded an architecture in which the dipyrromethene unit was coordinated to a boron(III) center within the macrocyclic cavity. Notably, the threaded BODIPY exhibited higher fluorescence quantum yields than the corresponding non-interlocked analogues. This strategy, therefore, provided a promising approach to developing novel, highly emissive dyes. The same group also developed a rotaxane architecture featuring threaded BODIPY units acting as stoppers.¹¹⁶ The resulting system exhibited high molar absorption coefficients and fluorescence quantum yields, as well as remarkable brightness.

Iminopyrrole self-assembly and dynamic supramolecular architectures

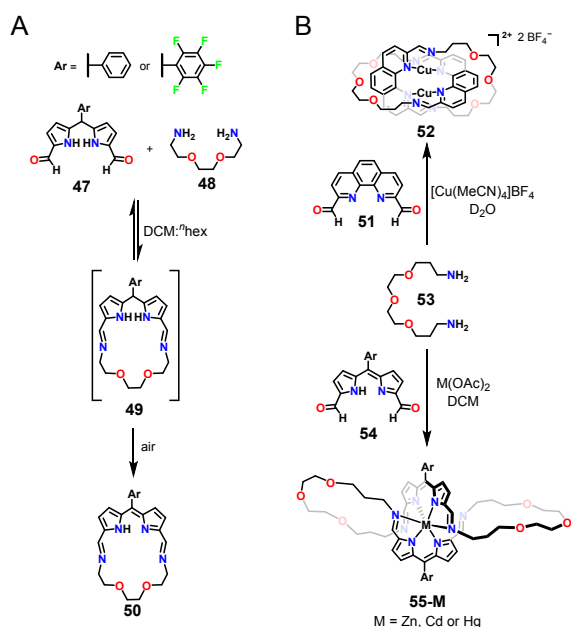
From the beginning, our research group has focused on two main areas that complement one another. These themes include the chemistry of mechanically interlocked molecules and the use of self-assembly to construct dynamic supramolecular systems. We are particularly interested in iminopyrrole ligands, an area actively explored by several research groups, including those of Sessler, Love, Fout, and others. What sets our work apart is the approach taken to constructing iminopyrrole architectures. Specifically, we rely on multicomponent reactions operating under thermodynamic control, in a manner conceptually related to the subcomponent self-assembly strategy.¹¹⁷

One of the first projects carried out in the group involved the synthesis of an original class of iminopyrrole macrocycles that combined structural features of crown ethers and porphyrins, termed crownphyrins.^{118,119} A related concept was proposed as early as the 1980s by Sessler; however, the use of



tripyrene building block prevented the formation of a conjugated heterocyclic brace within the macrocyclic ring.^{120,121} Later, Love and co-workers revisited this idea and reported a series of elegant transition-metal “Pac-Man” complexes based on macrocycles incorporating both porphyrinoid and crown ether elements.^{122,123}

Our approach relied on the use of a *meso*-pentafluorophenyl-substituted diformylpyrromethane precursor **47**, which, upon condensation with diamines bearing oligo(ethylene glycol) motifs **48**, formed macrocycles that underwent spontaneous oxidation of the intermediate crownphyrinogens **49** to the corresponding crownphyrins **50** (Scheme 17A).¹¹⁹



Scheme 17. A) Synthesis and reactivity of crownphyrins.¹¹⁹ B) Different reactivity of diformylphenanthroline **51** and diformylpyrromethene **54** in the metal-mediated self-assembly with **53**.^{117,119}

The hybrids turned out to be excellent receptors, as their cavities contain both hydrogen-bond donors and acceptors. As a result, they efficiently bind neutral molecules, such as water, and anions (after reduction), and act as versatile ligands.¹²⁴ By selecting macrocycles of appropriate size, monomeric mononuclear complexes **50-Pb**, dimeric “accordion-type” porphyrinoids,¹²⁵ as well as figure-eight dimers **55-Zn** exhibiting dynamic coordination behaviour in solution, were obtained (Figure 8). Reactions with Group 1 metal ions were found to engage the ether oxygen atoms only, leading to the formation of intriguing $M_2(H_2O)_2$ -bridged dimers **57-M** ($M = Na, K$).¹²⁴ Particularly interesting was the reactivity toward Pd(II), which enforced previously unobserved iminopyrrole tautomerism involving proton transfer to a carbon atom. Finally, larger macrocycles displayed Janus-type behaviour: Pd(II) was coordinated within the porphyrinoid core, while alkali metal cation occupied the crown-ether-like pocket.

Inspired by the work of Nitschke and co-workers,¹¹⁷ who demonstrated that the reaction between phenanthroline-2,9-dicarboxaldehyde **51** and 2,2'-(ethane-1,2-

diylbis(oxy))bis(ethan-1-amine) **53** in the presence of a Cu(I) afforded a template-stabilized [2]catenane **52**, we set out to investigate a similar transformation employing diformylpyrromethene **54** (Scheme 17B). Copper(I) proved unsuitable under our reaction conditions, as it was readily oxidized to Cu(II), reflecting the tendency of iminopyrrole ligands to stabilize higher oxidation states of metal ions. At this stage, a fundamental difference between the iminophenanthroline and iminopyrrole systems became apparent. In our case, the sole product formed was the figure-eight complex **55-M**.

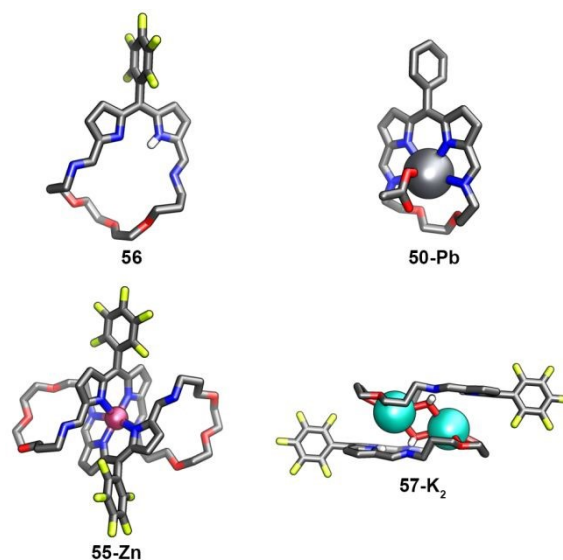


Figure 8. The X-ray molecular structure of **56**, **50-Pb**, **55-Zn**, and **57-K₂**.^{119,124}

Although we did not succeed at that time in obtaining a mechanically interlocked architecture, ongoing efforts are aimed at exploiting crownphyrin-based systems toward this goal. Importantly, the observed reactivity prompted us to further explore iminopyrrole ligands in subcomponent self-assembly processes. To evaluate the potential of iminopyrroles for constructing molecular systems of increasing complexity, we next examined their applicability in the assembly of capsular assemblies. Although purely organic cages based on iminopyrroles had already been reported at the time we initiated our studies,^{126–129} metal–organic assemblies derived from iminopyrroles remained largely unexplored.^{130,131}

In our initial approach, we investigated the reactions between 2,5-diformylpyrrole, tris(2-aminoethyl)amine (tren), and silver(I) salts.¹³² We found that the outcome strongly depended on the nature of the silver salt, more specifically on the coordinating properties of the counteranion. Two principal types of products were obtained: cascade-type cryptates **58-Ag₂F**, in which two silver ions were located at opposite poles of a [3+2] cage with a central fluoride (or chloride anion(s)), and so-called “plenates”, e.g., **58-Ag₃**, i.e., cages whose cavities were filled with a silver(I) cluster composed of three to five metal ions (Figure 9). Interestingly, theoretical studies have shown that the encapsulated silver clusters exhibit dynamic behaviour, with metal ions continuously rearranging their



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coordination environments and adopting multiple geometries within the cavity.¹³³ Subsequently, we found that the simple [3+2] cage can encapsulate two metal centers of the same type, such as Zn(II), Cd(II), or Hg(II) (**58-Hg₂**), as well as form heterobinuclear assemblies.¹³⁴ In these complexes, the cage framework stabilized very short metal–metal contacts, resulting in metallophilic interaction.

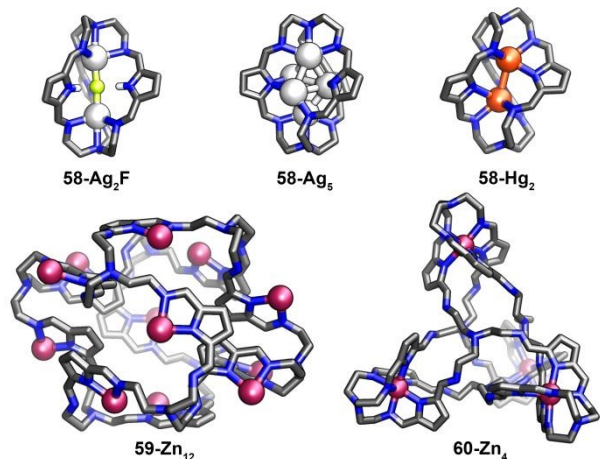


Figure 9. The X-ray molecular structures of metal-stabilized iminopyrrole cages **58-Ag₂F**, **58-Ag₅**, **58-Hg₂**, **59-Zn₁₂**, **60-Zn₄**.^{132,134,135}

An intriguing observation was made during studies of cages assembled in the presence of Zn(II).¹³⁵ In this case, the products formed by self-assembly were found to depend strongly on the solvent used. When the reaction was carried out in *n*-butanol, the expected [3+2] cage incorporating two Zn(II) ions was obtained. In contrast, performing the reaction in chloroform led to the formation of a much larger [12+8] system **59-Zn₁₂** stabilized by the coordination of twelve Zn(II) ions (Figure 9). The formation of the latter required partial tautomerization of the organic ligand from the diiminopyrrole form to an imino–aminoazafulvene in a tautomerism-coupled self-assembly (Scheme 18A).¹³⁶ The protons located on the amino groups of the latter were strongly engaged in hydrogen bonding, creating an extensive network of stabilizing interactions. Treatment with a competitive pyridine ligand induced transformation of this assembly into a cage **60-Zn₄**, in which four octahedral Zn(II) centres occupied the vertices of a tetrahedron.

Eventually, we sought to determine whether iminopyrroles could serve as building blocks for complex metallo-supramolecular architectures, including mechanically interlocked molecules such as knots and links. To this end, the replacement of 2,6-diformylpyridine by 2,5-diformylpyrrole in the reactions with bipyridyl-containing diamines developed by Stoddart and Trabolsi,^{137,138} carried out in the presence of Zn(II) or Cd(II), afforded [2]catenane **61-Zn₂**, trefoil knot **62-Cd₃**, and Borromean rings **63-Cd₆** incorporating diiminopyrrole motif (Figure 10).¹³⁹ Interestingly, the X-ray diffraction analysis

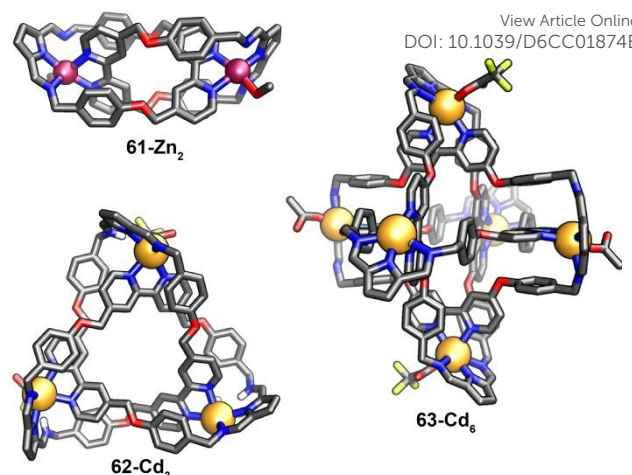
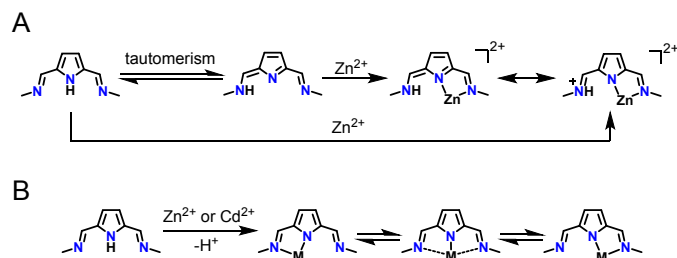


Figure 10. The X-ray molecular structures of **61-Zn₂**, **62-Cd₃**, **63-Cd₆**.¹³⁹

revealed that, contrary to the high symmetry anticipated based on the ¹H NMR spectra recorded at 300 K, the diiminopyrrole motif acted primarily as a bidentate, not a tridentate ligand within the coordination pocket of the assembly. This observation was readily rationalized by considering the intrinsic intramolecular dynamics of the assemblies, including the continuous cleavage and reforming of N–M bonds within diiminopyrrole coordination compound (Scheme 18B). Consequently, such dynamic behavior appeared to be quite a common feature of assemblies incorporating the iminopyrrole coordination motif.

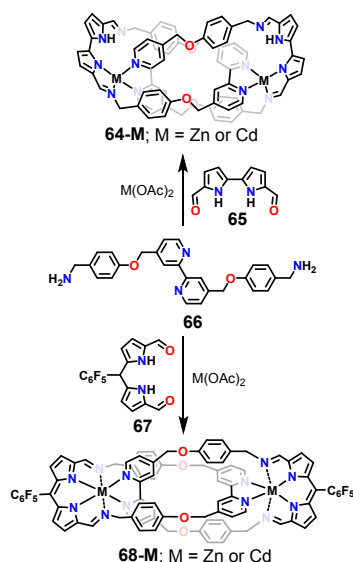


Scheme 18. A) Tautomerism in diiminopyrrole-based ligands.¹³⁶ B) Intramolecular dynamics of coordination compounds of diiminopyrroles.¹³⁹

Next, we explored the use of more structurally elaborate pyrrolic ligands in self-assembly processes, going beyond simple iminopyrroles to building blocks reminiscent of those employed in the construction of porphyrins and porphyrinoids.¹⁴⁰ In this context, we investigated the reactivity of two aldehydes – diformylbipyrrole **65** and diformyldipyrromethane **67** (Scheme 19). The latter is a direct precursor in the synthesis of porphyrins and numerous other pyrrolic macrocycles; upon oxidation, it forms dipyrin, often considered as a “half-porphyrin” subunit. In contrast, bipyrrole is found in the frameworks of contracted porphyrinoids such as corroles, as well as in expanded systems including sapphyrins.¹⁴¹ When these aldehydes were subjected to condensation with amine **66** in the presence of Zn(II) or Cd(II) salts, the only products obtained were metal-stabilised [2]catenanes, which we termed catenaphyrins. Interestingly, those two assemblies also demonstrated dynamic behaviour in

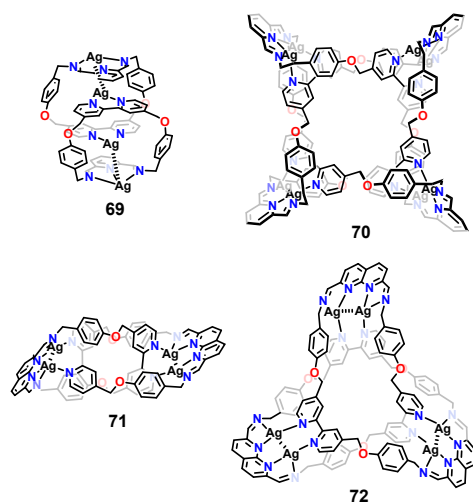


the solution. Notably, **68-M** exhibited interesting optical properties, with an emission centered at 580–620 nm. This exploratory study confirmed that more intricate pyrrole-based building blocks can be successfully employed in self-assembly processes, opening a route to the construction of architectures such as hypothetical knotaphyrins, i.e., molecules that blur the boundaries between molecular knots and expanded porphyrinoids.



Scheme 19. Self-assembly of [2]catenaphyrins **64-M** and **68-M**.¹⁴⁰

In parallel, we investigated self-assembly reactions leading to molecular knots and links stabilized by silver(I) ions. During these studies, the formation of an intriguing molecular motif, namely, a molecular tweezer **69**, was detected. The latter formed from the reaction of 2,2'-bipyridyl-based diamine with 2,6-diformylpyridine in the presence of silver(I) acetate (Figure 11). The structure consisted of two macrocyclic units incorporating a diiminopyridine motif, with dimerization driven by coordination of Ag(I)⋯Ag(I) pairs.



Scheme 20. The Ag(I)⋯Ag(I)-stabilized assemblies **69-72**.¹⁴²

Remarkably, a subtle modification of the amine component used in the self-assembly process enabled the selective formation of a Solomon link **70** stabilized by four Ag₂²⁺ units. Further investigations revealed that this silver cluster can serve as an effective template, enabling the formation of [2]catenane **71** and trefoil knot **72** when 2,7-diformyl-1,8-naphthyridine was used as the carbonyl component in the self-assembly reaction (Scheme 20). Although compounds **69-72** did not incorporate pyrrole-based ligands, they represent an extension of our earlier work on silver cluster templating, with potential applicability to pyrrolic and other assemblies.

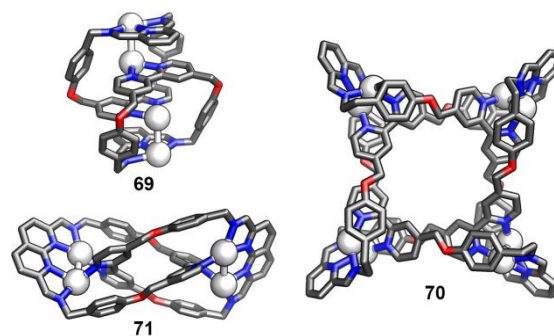


Figure 11. The X-ray molecular structures of **69**, **70**, and **71**.¹⁴²

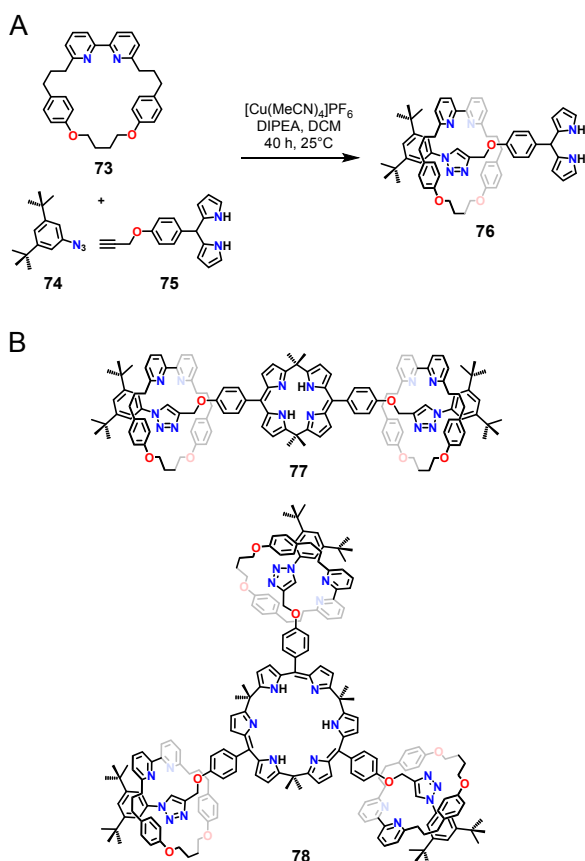
Calixphyrin-based catenanes and rotaxanes

In search of a general strategy for the synthesis of diverse mechanically interlocked porphyrinoids, we reasoned that access to a versatile and fundamental building block – namely a [2]rotaxane stoppered with a dipyrromethane such as **77** – would enable the preparation of a broad family of such systems.¹⁴³ This compound was obtained using a macrocyclic bipyridine **73** reported by Goldup and co-workers.¹⁴⁴ In a reaction between azide **74** and a *meso*-4-alkynyl-substituted dipyrromethane **75** under copper(I)-catalyzed active-template conditions,^{145–147} the desired mechanically interlocked dipyrromethane **76** was formed in high yield (83%) (Scheme 21A). Interestingly, the dipyrromethane stopper was sufficiently large to prevent dethreading under conditions of porphyrinoid condensation. Furthermore, the successful synthesis of a [2]rotaxane featuring a long thread terminated with a dipyrromethane at both ends enabled the subsequent design of the corresponding catenanes.

Rotaxane **76** proved to be an excellent building block for the synthesis of porphyrinoids. Upon reaction with pentafluorobenzaldehyde under Lindsey condensation conditions, formation of a [3]rotaxane was observed, featuring an unsubstituted porphyrin unit positioned at the center of the axle. Having confirmed the utility of this building block in porphyrinoid-forming reactions, we next investigated its reaction with acetone. This transformation afforded two products: a [3]rotaxane **77** and a Y-shaped [4]rotaxane **78**, incorporating calix[4]phyrin and calix[6]phyrin units, respectively (Scheme 21B). The core calix[*n*]phyrins represent hybrid macrocycles that bridge nonconjugated calix[*n*]pyrroles and fully conjugated porphyrins, owing to the presence of *meso*-sp³ carbon bridges



separating the dipyrromethene units within the macrocycle.^{148,149} As a consequence, calix[4]phyrins adopt a roof-like conformation in the solid state, whereas in solution they exhibit dynamic behavior involving interconversion between V-shaped stereoisomers, leading to time-averaged positions of the methyl groups (Figure 12A).¹⁵⁰ Notably, this conformational switching is retained upon incorporation of the calix[4]phyrin unit into the axle of a rotaxane, thereby inducing motion of the entire mechanically interlocked molecule (Figure 12B).¹⁴³ Owing to its resemblance to the motion of bird or butterfly wings during flight, we termed this dynamic behaviour as fluttering. Interestingly, the fluttering motion was suppressed when the calix[4]phyrin unit was incorporated into a [2]catenane, most likely due to an insufficient cavity size that restricts the conformational flexibility. Furthermore, the [3]rotaxane, by virtue of its molecular design, exhibited multimodal motion in response to external stimuli. In addition to fluttering – which can be reversibly switched on and off by (de)protonation of the calixphyrin core – the system also undergoes macrocycle shuttling along the thread and rotational motion of the macrocyclic component. Further strategies to control the fluttering motion are currently being explored in our laboratory.



Scheme 21. A) Synthesis of a dipyrromethane-stoppered rotaxane **76** and B) its condensation products – [3]rotaxane **77** and [4]rotaxane **78**.¹⁴³

The developed strategy for the synthesis of [2]rotaxanes bearing dipyrromethene units at both termini enabled the design of a series of precursors for the construction of [n]catenanes.¹⁵¹ Under optimized conditions, substrates differing in the length and flexibility of the linear component were expected to undergo condensation to

form interlocked products. It was ultimately found that the outcome strongly depended on the thread's geometric features. Rotaxanes with sufficiently long and flexible linear components underwent intramolecular self-condensation in the presence of acetone and a catalyst, affording [2]catenanes. In contrast, shorter and more rigid [2]rotaxanes favored intermolecular condensation, leading to the formation of higher-order [3]- and [4]catenanes. These studies demonstrated that the structure of the starting materials plays a decisive role in determining the topology of the final product. Consequently, this methodology is expected to be extendable to the synthesis of higher-order rotaxanes and catenanes based on diverse scaffolds, not limited to the calixphyrin motif.

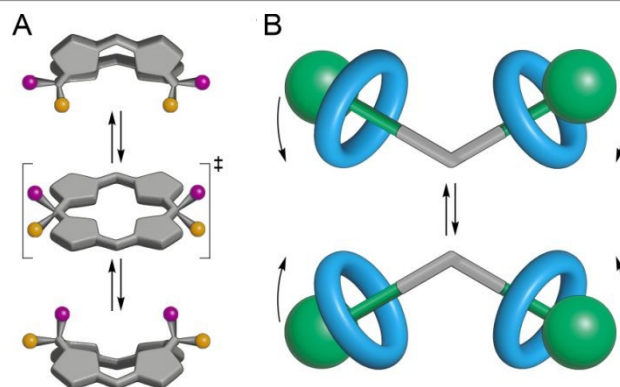


Figure 12 A) The toggling in calix[4]phyrin.^{149,150} B) Fluttering in calix[4]phyrin incorporating [3]rotaxane.¹⁴³

Conclusions

The chemistry at the interface between pyrrolic macrocycles and mechanically interlocked molecules, although initiated several decades ago, continues to evolve. While a number of structurally and functionally intriguing architectures have emerged from combining these two classes of molecular systems, the field remains dominated by porphyrin-based examples, with porphyrinoids or their fragments incorporated into MIMs only occasionally. Yet, the potential of such systems appears far from fully realized. Embedding porphyrinoid within the constrained environment of MIMs offers a powerful opportunity to introduce new dimensions of functionality, including controlled modulation of conjugation pathways, the emergence of spatially differentiated magnetic domains, and reversible switching between multiple electronic states. Moreover, mechanically interlocked architectures provide a unique platform for probing how the fundamental properties of pyrrolic macrocycles alter under topological constraints. It is anticipated that deeper exploration of this interface will not only enrich the chemistry of MIMs but also open new conceptual avenues in porphyrinoid chemistry, where topology and electronic structure can be deliberately intertwined.

Author contributions

† These Authors contributed equally.



Data availability

No primary research results have been included, and no new data were generated or analysed as part of this feature article.

Conflicts of interest

There are no conflicts to declare.

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References

- C. J. Brunts and J. F. Stoddart, *The Nature of the Mechanical Bond: From Molecules to Machines*, Wiley, 1st edn., 2016.
- S. Erbas-Cakmak, D. A. Leigh, C. T. McTernan and A. L. Nussbaumer, *Chem. Rev.*, 2015, **115**, 10081–10206.
- J. M. Abendroth, O. S. Bushuyev, P. S. Weiss and C. J. Barrett, *ACS Nano*, 2015, **9**, 7746–7768.
- V. Amendola, L. Fabbrizzi, C. Mangano and P. Pallavicini, *Acc. Chem. Res.*, 2001, **34**, 488–493.
- M. Russew and S. Hecht, *Adv. Mater.*, 2010, **22**, 3348–3360.
- R. G. E. Coumans, J. A. A. W. Elemans, A. E. Rowan and R. J. M. Nolte, *Chem. Eur. J.*, 2013, **19**, 7758–7770.
- B. L. Feringa, N. Koumura, R. A. van Delden and M. K. J. ter Wiel, *Appl. Phys. A*, 2002, **75**, 301–308.
- T. L. Mako, J. M. Racicot and M. Levine, *Chem. Rev.*, 2019, **119**, 322–477.
- J. T. Wilmore and P. D. Beer, *Adv. Mater.*, 2024, **36**, 2309098.
- L. F. Hart, J. E. Hertzog, P. M. Rauscher, B. W. Rawe, M. M. Tranquilli and S. J. Rowan, *Nat. Rev. Mater.*, 2021, **6**, 508–530.
- S. Mena-Hernando and E. M. Pérez, *Chem. Soc. Rev.*, 2019, **48**, 5016–5032.
- R. A. Grzelczak, A. Władyczyn, A. Białońska, Ł. John and B. Szyszko, *Chem. Commun.*, 2023, **59**, 7579–7582.
- J. A. Faiz, V. Heitz and J.-P. Sauvage, *Chem. Soc. Rev.*, 2009, **38**, 422–442.
- A. Harriman and J.-P. Sauvage, *Chem. Soc. Rev.*, 1996, **25**, 41–48.
- J.-C. Chambron, J.-P. Collin, J.-O. Dalbavie, C. O. Dietrich-Buchecker, V. Heitz, F. Odobel, N. Solladié and J.-P. Sauvage, *Coord. Chem. Rev.*, 1998, **178–180**, 1299–1312.
- L. Flamigni, V. Heitz and J.-P. Sauvage, in *Non-Covalent Multi-Porphyrin Assemblies*, ed. E. Alessio, Springer, Berlin, Heidelberg, 2006, 217–261.
- J.-C. Chambron, V. Heitz and J.-P. Sauvage, *J. Chem. Soc., Chem. Commun.*, 1992, 1131–1133.
- J.-C. Chambron, V. Heitz and J.-P. Sauvage, *J. Am. Chem. Soc.*, 1993, **115**, 12378–12384.
- P. R. Ashton, M. R. Johnston, J. F. Stoddart, M. S. Tolley and J. W. Wheeler, *J. Chem. Soc., Chem. Commun.*, 1992, 1128–1131.
- Y. Wu, M. Frasconi, W.-G. Liu, R. M. Young, W. A. Goddard, M. R. Wasielewski and J. F. Stoddart, *J. Am. Chem. Soc.*, 2020, **142**, 11835–11846.
- M. J. Gunter, N. Bampos, K. D. Johnstone and J. K. M. Sanders, *New J. Chem.*, 2001, **25**, 166–173.
- K. D. Johnstone, N. Bampos, J. K. M. Sanders and M. J. Gunter, *Chem. Commun.*, 2003, 1396–1397.
- K. M. Mullen, K. D. Johnstone, D. Nath, N. Bampos, J. K. M. Sanders and M. J. Gunter, *Org. Biomol. Chem.*, 2008, **7**, 293–303.
- K. M. Mullen, K. D. Johnstone, M. Webb, N. Bampos, J. K. M. Sanders and M. J. Gunter, *Org. Biomol. Chem.*, 2008, **6**, 278–286.
- C. A. Hunter and L. D. Sarson, *Angew. Chem. Int. Ed.*, 1994, **33**, 2313–2316.
- C. A. Hunter, C. M. R. Low, M. J. Packer, S. E. Spey, J. G. Vinter, M. O. Vysotsky and C. Zonta, *Angew. Chem. Int. Ed.*, 2001, **40**, 2678–2682.
- M. J. Langton, J. D. Matichak, A. L. Thompson and H. L. Anderson, *Chem. Sci.*, 2011, **2**, 1897–1901.
- D. R. Kohn, L. D. Movsisyan, A. L. Thompson and H. L. Anderson, *Org. Lett.*, 2017, **19**, 348–351.
- J.-P. Sauvage, J.-P. Collin, J. A. Faiz, J. Frey, V. Heitz and C. Tock, *J. Porphyr. Phthalocyanines*, 2008, **12**, 881–905.
- B. Szyszko and L. Latos-Grażyński, *Chem. Soc. Rev.*, 2015, **44**, 3588–3616.
- P. J. Chmielewski and L. Latos-Grażyński, *Coord. Chem. Rev.*, 2005, **249**, 2510–2533.
- T. M. Krygowski and M. K. Cyranski, Eds., *Aromaticity in Heterocyclic Compounds*, Springer Berlin Heidelberg, Berlin, Heidelberg, 2009, vol. 19.
- M. Stępień, N. Sprutta and L. Latos-Grażyński, *Angew. Chem. Int. Ed.*, 2011, **50**, 4288–4340.
- B. Szyszko, M. J. Białek, E. Pacholska-Dudziak and L. Latos-Grażyński, *Chem. Rev.*, 2017, **117**, 2839–2909.
- B. Szyszko, M. Przewoźnik, M. J. Białek, A. Białońska, P. J. Chmielewski and L. Latos-Grażyński, *Chem. Eur. J.*, 2020, **26**, 8555–8566.
- B. Szyszko, E. Pacholska-Dudziak and L. Latos-Grażyński, *J. Org. Chem.*, 2013, **78**, 5090–5095.
- P. Krzyszowska, A. Burska-Jabłońska, M. Oberski, M. J. Białek, L. Latos-Grażyński and K. Hurej, *Angew. Chem. Int. Ed.*, 2026, **65**, e25506.
- S. Sanfui, M. Usman, S. Sarkar, S. Pramanik, E. Garribba and S. P. Rath, *Inorg. Chem.*, 2022, **61**, 8419–8430.
- M. T. Green, J. H. Dawson and H. B. Gray, *Science*, 2004, **304**, 1653–1656.
- W. J. Song, M. S. Seo, S. DeBeer George, T. Ohta, R. Song, M.-J. Kang, T. Tosha, T. Kitagawa, E. I. Solomon and W. Nam, *J. Am. Chem. Soc.*, 2007, **129**, 1268–1277.
- M. J. Białek, K. Hurej, H. Furuta and L. Latos-Grażyński, *Chem. Soc. Rev.*, 2023, **52**, 2082–2144.
- T. Higashino, A. Kumagai, S. Sakaki and H. Imahori, *Chem. Sci.*, 2018, **9**, 7528–7539.
- L. Raehm, J.-M. Kern and J.-P. Sauvage, *Chem. Eur. J.*, 1999, **5**, 3310–3317.
- R. A. Bissell, E. Córdova, A. E. Kaifer and J. F. Stoddart, *Nature*, 1994, **369**, 133–137.
- M. J. Power, D. T. J. Morris, I. J. Vitorica-Yrezabal and D. A. Leigh, *J. Am. Chem. Soc.*, 2023, **145**, 8593–8599.
- C. M. Storey, M. R. Gyton, R. E. Andrew and A. B. Chaplin, *Angew. Chem. Int. Ed.*, 2018, **57**, 12003–12006.
- M.-J. Blanco, J.-C. Chambron, V. Heitz and J.-P. Sauvage, *Org. Lett.*, 2000, **2**, 3051–3054.
- J. Wu, Fang, W.-Y. Lu, J.-L. Hou, C. Li, Z.-Q. Wu, X.-K. Jiang, Z.-T. Li and Y.-H. Yu, *J. Org. Chem.*, 2007, **72**, 2897–2905.
- N. Solladié, J.-C. Chambron, C. O. Dietrich-Buchecker and J.-P. Sauvage, *Angew. Chem. Int. Ed.*, 1996, **35**, 906–909.
- N. Solladié, J.-C. Chambron and J.-P. Sauvage, *J. Am. Chem. Soc.*, 1999, **121**, 3684–3692.
- K. Chichak, M. C. Walsh and N. R. Branda, *Chem. Commun.*, 2000, 847–848.
- M. Asakawa, T. Ikeda, N. Yui and T. Shimizu, *Chem. Lett.*, 2002, **31**, 174–175.



- 53 C. Roche, J.-P. Sauvage, A. Sour and N. L. Strutt, *New J. Chem.*, 2011, **35**, 2820–2825.
- 54 P. Martinez-Bulit, B. H. Wilson and S. J. Loeb, *Org. Biomol. Chem.*, 2020, **18**, 4395–4400.
- 55 L. D. Movsisyan, M. Franz, F. Hampel, A. L. Thompson, R. R. Tykwinski and H. L. Anderson, *J. Am. Chem. Soc.*, 2016, **138**, 1366–1376.
- 56 J. Pruchyathamkorn, W. J. Kendrick, A. T. Frawley, A. Mattioni, F. Caycedo-Soler, S. F. Huelga, M. B. Plenio and H. L. Anderson, *Angew. Chem. Int. Ed.*, 2020, **59**, 16455–16458.
- 57 V. Heitz, S. Chardon-Noblat and J.-P. Sauvage, *Tetrahedron Lett.*, 1991, **32**, 197–198.
- 58 J. C. Chambron, A. Harriman, V. Heitz and J. P. Sauvage, *J. Am. Chem. Soc.*, 1993, **115**, 6109–6114.
- 59 J. C. Chambron, A. Harriman, V. Heitz and J. P. Sauvage, *J. Am. Chem. Soc.*, 1993, **115**, 7419–7425.
- 60 M. Linke, J.-C. Chambron, V. Heitz and J.-P. Sauvage, *J. Am. Chem. Soc.*, 1997, **119**, 11329–11330.
- 61 M. Linke, J.-C. Chambron, V. Heitz, J.-P. Sauvage, S. Encinas, F. Barigelletti and L. Flamigni, *J. Am. Chem. Soc.*, 2000, **122**, 11834–11844.
- 62 L. Flamigni, A. M. Talarico, S. Serroni, F. Puntoriero, M. J. Gunter, M. R. Johnston and T. P. Jeynes, *Chem. Eur. J.*, 2003, **9**, 2649–2659.
- 63 D. Gust, T. A. Moore and A. L. Moore, *Acc. Chem. Res.*, 2001, **34**, 40–48.
- 64 D. Kuciauskas, P. A. Liddell, S. Lin, S. G. Stone, A. L. Moore, T. A. Moore and D. Gust, *J. Phys. Chem. B*, 2000, **104**, 4307–4321.
- 65 D. I. Schuster, K. Li, D. M. Guldi and J. Ramey, *Org. Lett.*, 2004, **6**, 1919–1922.
- 66 A. S. D. Sandanayaka, N. Watanabe, K.-I. Ikeshita, Y. Araki, N. Kihara, Y. Furusho, O. Ito and T. Takata, *J. Phys. Chem. B*, 2005, **109**, 2516–2525.
- 67 N. Watanabe, N. Kihara, Y. Furusho, T. Takata, Y. Araki and O. Ito, *Angew. Chem. Int. Ed.*, 2003, **42**, 681–683.
- 68 K. Li, D. I. Schuster, D. M. Guldi, M. Á. Herranz and L. Echegoyen, *J. Am. Chem. Soc.*, 2004, **126**, 3388–3389.
- 69 J. D. Megiatto, D. I. Schuster, S. Abwandner, G. De Miguel and D. M. Guldi, *J. Am. Chem. Soc.*, 2010, **132**, 3847–3861.
- 70 F. Schwer, S. Zank, M. Freiburger, F. M. Steudel, N. Geue, L. Ye, P. E. Barran, T. Drewello, D. M. Guldi and M. Von Delius, *Angew. Chem.*, 2025, **137**, e202413404.
- 71 M. J. Gunter, T. P. Jeynes and P. Turner, *Eur. J. Org. Chem.*, 2004, **2004**, 193–208.
- 72 D. P. Cormode, M. G. B. Drew, R. Jagessar and P. D. Beer, *Dalton Trans.*, 2008, 6732–6741.
- 73 J.-P. Collin, F. Durolo, J. Frey, V. Heitz, J.-P. Sauvage, C. Tock and Y. Trolez, *Chem. Commun.*, 2009, 1706–1708.
- 74 J.-S. Marois, K. Cantin, A. Desmarais and J.-F. Morin, *Org. Lett.*, 2008, **10**, 33–36.
- 75 J. Frey, C. Tock, J.-P. Collin, V. Heitz and J.-P. Sauvage, *J. Am. Chem. Soc.*, 2008, **130**, 4592–4593.
- 76 A. Brown and P. D. Beer, *Dalton Trans.*, 2012, **41**, 118–129.
- 77 Y. Cheong Tse, R. Hein, E. J. Mitchell, Z. Zhang and P. D. Beer, *Chem. Eur. J.*, 2021, **27**, 14550–14559.
- 78 A. Brown, M. J. Langton, N. L. Kilah, A. L. Thompson and P. D. Beer, *Chem. Eur. J.*, 2015, **21**, 17664–17675.
- 79 M. Linke, J.-C. Chambron, V. Heitz, J.-P. Sauvage and V. Semetey, *Chem. Commun.*, 1998, 2469–2470.
- 80 J.-P. Collin, J. Frey, V. Heitz, J.-P. Sauvage, C. Tock and L. Allouche, *J. Am. Chem. Soc.*, 2009, **131**, 5609–5620.
- 81 J. T. Wilmore, Y. Cheong Tse, A. Docker, C. Whitehead, C. K. Williams and P. D. Beer, *Chem. Eur. J.*, 2023, **29**, e202300608.
- 82 Y. Oka, H. Masai and J. Terao, *Angew. Chem. Int. Ed.*, 2023, **62**, e202217002.
- 83 P. Thordarson, E. J. A. Bijsterveld, A. E. Rowan and R. J. M. Nolte, *Nature*, 2003, **424**, 915–918.
- 84 R. G. E. Coumans, J. A. A. W. Elemans, R. J. M. Nolte and A. E. Rowan, *Proc. Natl. Acad. Sci. U.S.A.*, 2006, **103**, 19647–19651.
- 85 A. B. C. Deutman, C. Monnereau, J. A. A. W. Elemans, G. Ercolani, R. J. M. Nolte and A. E. Rowan, *Science*, 2008, **322**, 1668–1671.
- 86 C. Monnereau, P. H. Ramos, A. B. C. Deutman, J. A. A. W. Elemans, R. J. M. Nolte and A. E. Rowan, *J. Am. Chem. Soc.*, 2010, **132**, 1529–1531.
- 87 R. J. M. Nolte and J. A. A. W. Elemans, *Chem. Eur. J.*, 2024, **30**, e202304230.
- 88 A. F. P. Alcântara, L. A. Fontana, V. H. Rigolin, Y. F. S. Andrade, M. A. Ribeiro, W. P. Barros, C. Ornelas and J. D. Megiatto, *Angew. Chem. Int. Ed.*, 2018, **57**, 8979–8983.
- 89 L. A. Fontana, M. P. Almeida, A. F. P. Alcântara, V. H. Rigolin, M. A. Ribeiro, W. P. Barros and J. D. Megiatto, *Nat. Commun.*, 2020, **11**, 6370.
- 90 Y. Miyazaki, C. Kahlfuss, A. Ogawa, T. Matsumoto, J. A. Wytko, K. Oohora, T. Hayashi and J. Weiss, *Chem. Eur. J.*, 2017, **23**, 13579–13582.
- 91 A. Baeyer, *Ber. Dtsch. Chem. Ges.*, 1886, **19**, 2184–2185.
- 92 P. A. Gale, J. L. Sessler, V. Král and V. Lynch, *J. Am. Chem. Soc.*, 1996, **118**, 5140–5141.
- 93 V. Valderrey, E. C. Escudero-Adán and P. Ballester, *J. Am. Chem. Soc.*, 2012, **134**, 10733–10736.
- 94 J. R. Romero, G. Aragay and P. Ballester, *Chem. Sci.*, 2017, **8**, 491–498.
- 95 R. Molina-Muriel, J. R. Romero, Y. Li, G. Aragay and P. Ballester, *Org. Biomol. Chem.*, 2021, **19**, 9986–9995.
- 96 Y. Li, R. Molina-Muriel, G. Aragay and P. Ballester, *Org. Chem. Front.*, 2024, **11**, 5374–5384.
- 97 D. Memis, N. Bektas and A. Aydogan, *ACS Appl. Polym. Mater.*, 2024, **6**, 4339–4348.
- 98 D. T. Gryko and M. Tasiar, in *Fundamentals of Porphyrin Chemistry*, John Wiley & Sons, Ltd, 2022, pp. 349–383.
- 99 C. D. Natale, C. P. Gros and R. Paolesse, *Chem. Soc. Rev.*, 2022, **51**, 1277–1335.
- 100 T. H. Ngo, J. Labuta, G. N. Lim, W. A. Webre, F. D'Souza, P. A. Karr, J. E. M. Lewis, J. P. Hill, K. Ariga and S. M. Goldup, *Chem. Sci.*, 2017, **8**, 6679–6685.
- 101 T. H. Ngo, J. E. M. Lewis, D. T. Payne, F. D'Souza, J. P. Hill, K. Ariga, G. Yoshikawa and S. M. Goldup, *Chem. Commun.*, 2020, **56**, 7447–7450.
- 102 D. Seidel, V. Lynch and J. L. Sessler, *Angew. Chem. Int. Ed.*, 2002, **41**, 1422–1425.
- 103 T. Izumi, S. Kobashi, K. Takimiya, Y. Aso and T. Otsubo, *J. Am. Chem. Soc.*, 2003, **125**, 5286–5287.
- 104 P. Bäuerle, J. Becher, J. Lau and P. Mark, in *Electronic Materials: The Oligomer Approach*, John Wiley & Sons, Ltd, 1998, pp. 105–233.
- 105 M. Ammann, A. Rang, C. A. Schalley and P. Bäuerle, *Eur. J. Org. Chem.*, 2006, **2006**, 1940–1948.
- 106 P. Bäuerle, M. Ammann, M. Wilde, G. Götz, E. Mena-Osteritz, A. Rang and C. A. Schalley, *Angew. Chem. Int. Ed.*, 2007, **46**, 363–368.
- 107 G. Götz, X. Zhu, A. Mishra, J. Segura, E. Mena-Osteritz and P. Bäuerle, *Chem. Eur. J.*, 2015, **21**, 7193–7210.
- 108 Y. Sui, Z. Wang, J. Hao, S. Wu, L. Ren, Z. Zeng, J. Wu and Y. Ni, *Nat. Synth.*, 2025, **5**, 230–239.
- 109 A. Andrievsky, F. Ahuis, J. L. Sessler, F. Vögtle, D. Gudat and M. Moini, *J. Am. Chem. Soc.*, 1998, **120**, 9712–9713.
- 110 B. Nisanci, S. Sahinoglu, E. Tuner, M. Arik, İ. Kani, A. Dastan and Ö. A. Bozdemir, *Chem. Commun.*, 2017, **53**, 12418–12421.
- 111 N. Sakamoto, C. Ikeda and T. Nabeshima, *Chem. Commun.*, 2010, **46**, 6732–6734.
- 112 T. Nakamura, G. Yamaguchi and T. Nabeshima, *Angew. Chem. Int. Ed.*, 2016, **55**, 9606–9609.



- 113 N. Yesilgul, O. Seven, R. Guliyev and E. U. Akkaya, *J. Org. Chem.*, 2018, **83**, 13228–13232.
- 114 M. Hicguet, L. Verrieux, O. Mongin, T. Roisnel, F. Berrée, A. Fihey, B. Le Guennic and Y. Trolez, *Angew. Chem. Int. Ed.*, 2024, **63**, e202318297.
- 115 M. Hicguet, O. Mongin, Y. R. Leroux, T. Roisnel, F. Berrée and Y. Trolez, *ChemistryOpen*, 2024, **13**, e202400196.
- 116 M. Hicguet, O. Mongin, Y. R. Leroux, F. Berrée and Y. Trolez, *Org. Lett.*, 2025, **27**, 12217–12222.
- 117 M. Hutin, C. A. Schalley, G. Bernardinelli and J. R. Nitschke, *Chem. Eur. J.*, 2006, **12**, 4069–4076.
- 118 M. Matviyishyn and B. Szyszko, *Beilstein J. Org. Chem.*, 2023, **19**, 1630–1650.
- 119 M. Matviyishyn, A. Białońska and B. Szyszko, *Angew. Chem. Int. Ed.*, 2022, **61**, e202211671.
- 120 J. L. Sessler and A. K. Burrell, in *Macrocycles*, Springer Berlin Heidelberg, Berlin, Heidelberg, 1992, vol. 161, pp. 177–273.
- 121 J. L. Sessler, M. R. Johnson, V. Lynch and T. Murai, *J. Coord. Chem.*, 1988, **18**, 99–104.
- 122 J. W. Leeland, F. J. White and J. B. Love, *Chem. Commun.*, 2011, **47**, 4132–4134.
- 123 J. W. Leeland, C. Finn, B. Escuyer, H. Kawaguchi, G. S. Nichol, A. M. Z. Slawin and J. B. Love, *Dalton Trans.*, 2012, **41**, 13815–13831.
- 124 M. Matviyishyn, K. A. Konieczny, B. Trzaskowski and B. Szyszko, *Chem. Eur. J.*, 2024, **30**, e202402932.
- 125 W. A. Reiter, A. Gerges, S. Lee, T. Deffo, T. Clifford, A. Danby and K. Bowman-James, *Coord. Chem. Rev.*, 1998, **174**, 343–359.
- 126 D. Chen and A. E. Martell, *Tetrahedron*, 1991, **47**, 6895–6902.
- 127 H. J. Han, J. H. Oh, J. L. Sessler and S. K. Kim, *Chem. Commun.*, 2019, **55**, 10876–10879.
- 128 J. H. Oh, J. H. Kim, D. S. Kim, H. J. Han, V. M. Lynch, J. L. Sessler and S. K. Kim, *Org. Lett.*, 2019, **21**, 4336–4339.
- 129 O. D. Fox, T. D. Rolls, P. D. Beer and M. G. B. Drew, *Chem. Commun.*, 2001, 1632–1633.
- 130 L. Qin, J. Nelson and M. McCann, *J. Inorg. Biochem.*, 1993, **51**, 633–639.
- 131 Q. Lu, V. McKee and J. Nelson, *J. Chem. Soc., Chem. Commun.*, 1994, 649–651.
- 132 A. Sarwa, A. Białońska, M. Garbicz and B. Szyszko, *Chem. Eur. J.*, 2023, **29**, e202203850.
- 133 B. Trzaskowski, J. P. Martínez, A. Sarwa, B. Szyszko and W. A. Goddard, *J. Phys. Chem. A*, 2024, **128**, 3339–3350.
- 134 A. Sarwa, P. Krajewski, B. Trzaskowski, J. P. Perdek, M. Siczek and B. Szyszko, *Inorg. Chem.*, 2026, **65**, 5280–5285.
- 135 J. Sukiennik, A. Sarwa, J. P. Perdek, M. Siczek and B. Szyszko, *Chem. Eur. J.*, 2025, **31**, e02714.
- 136 M. Giri, D. Sahoo, B. P. Samantray, P. P. Sahoo, S. Mishra and T. Guchhait, *J. Mol. Struct.*, 2024, **1308**, 138029.
- 137 K. S. Chichak, S. J. Cantrill, A. R. Pease, S.-H. Chiu, G. W. V. Cave, J. L. Atwood and J. F. Stoddart, *Science*, 2004, **304**, 1308–1312.
- 138 T. Prakasam, M. Lusi, M. Elhabiri, C. Platas-Iglesias, J.-C. Olsen, Z. Asfari, S. Cianférani-Sanglier, F. Debaene, L. J. Charbonnière and A. Trabolsi, *Angew. Chem. Int. Ed.*, 2013, **52**, 9956–9960.
- 139 A. Sarwa, A. Białońska, M. Sobieraj, J. P. Martínez, B. Trzaskowski and B. Szyszko, *Angew. Chem. Int. Ed.*, 2024, **63**, e202316489.
- 140 A. Sarwa, M. Matviyishyn, J. P. Perdek, B. Trzaskowski, D. Kulesza, E. Zych and B. Szyszko, *Chem. Commun.*, 2026, **62**, 148–151.
- 141 V. J. Bauer, D. L. J. Clive, D. Dolphin, J. B. Paine, F. L. Harris, M. M. King, J. Loder, S. W. C. Wang and R. B. Woodward, *J. Am. Chem. Soc.*, 1983, **105**, 6429–6436.
- 142 A. Sarwa, A. Khmara, K. A. Konieczny, D. Kulesza, E. Zych, B. Trzaskowski and B. Szyszko, *Angew. Chem. Int. Ed.*, 2025, **64**, e202423962.
- 143 R. A. Grzelczak, T. Basak, B. Trzaskowski, V. Kinzhyballo and B. Szyszko, *Angew. Chem. Int. Ed.*, 2025, **64**, e202413579.
- 144 J. E. M. Lewis, R. J. Bordoli, M. Denis, C. J. Fletcher, M. Galli, E. A. Neal, E. M. Rochette and S. M. Goldup, *Chem. Sci.*, 2016, **7**, 3154–3161.
- 145 V. Aucagne, J. Berná, J. D. Crowley, S. M. Goldup, K. D. Hänni, D. A. Leigh, P. J. Lusby, V. E. Ronaldson, A. M. Z. Slawin, A. Viterisi and D. B. Walker, *J. Am. Chem. Soc.*, 2007, **129**, 11950–11963.
- 146 J. D. Crowley, S. M. Goldup, A.-L. Lee, D. A. Leigh and R. T. McBurney, *Chem. Soc. Rev.*, 2009, **38**, 1530–1541.
- 147 S. M. Goldup, D. A. Leigh, T. Long, P. R. McGonigal, M. D. Symes and J. Wu, *J. Am. Chem. Soc.*, 2009, **131**, 15924–15929.
- 148 C. Bucher, R. S. Zimmerman, V. Lynch, V. Král and J. L. Sessler, *J. Am. Chem. Soc.*, 2001, **123**, 2099–2100.
- 149 V. Král, J. L. Sessler, R. S. Zimmerman, D. Seidel, V. Lynch and B. Andrioletti, *Angew. Chem. Int. Ed.*, 2000, **39**, 1055–1058.
- 150 B. Dolenský, J. Kroulík, V. Král, J. L. Sessler, H. Dvořáková, P. Bouř, M. Bernátková, C. Bucher and V. Lynch, *J. Am. Chem. Soc.*, 2004, **126**, 13714–13722.
- 151 R. A. Grzelczak, J. P. Perdek, M. Siczek, P. J. Chmielewski and B. Szyszko, *Org. Lett.*, 2025, **27**, 6310–6315.



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Data availability

No primary research results have been included, and no new data were generated or analysed as part of this feature article.

