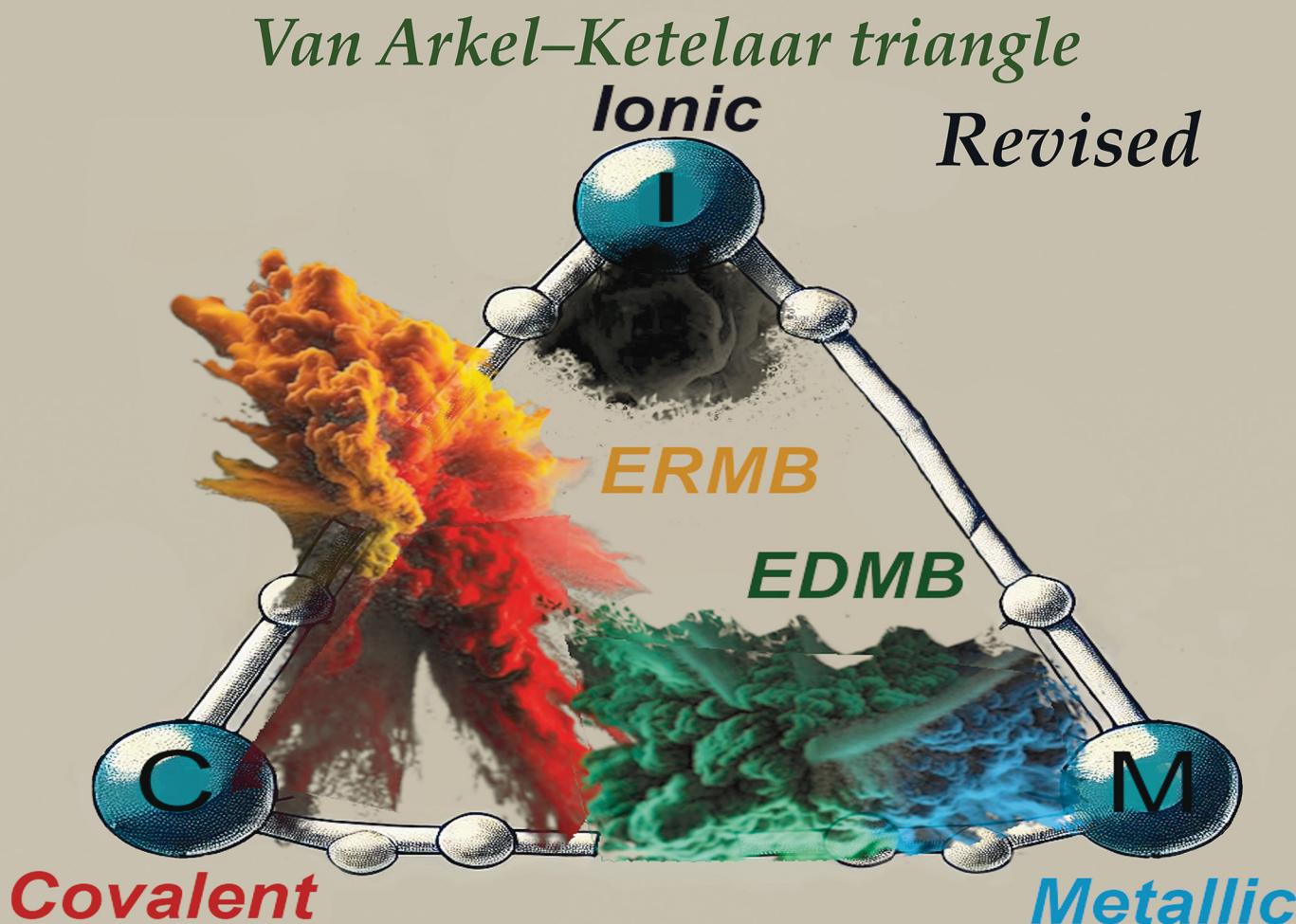


Journal of Materials Chemistry C

Materials for optical, magnetic and electronic devices

rsc.li/materials-c



ISSN 2050-7526

PAPER

H. H. Osman, F. J. Manjón *et al.*
A unified theory of electron-rich and electron-deficient
multicenter bonds in molecules and solids: a change of
paradigms



Cite this: *J. Mater. Chem. C*, 2025,
13, 3774

A unified theory of electron-rich and electron-deficient multicenter bonds in molecules and solids: a change of paradigms†

H. H. Osman *^{abc} P. Rodríguez-Hernández, ^d A. Muñoz ^d and F. J. Manjón *^a

Here, we propose a multicenter bond theory that addresses the origin and mechanisms behind the formation of electron-rich multicenter bonds (ERMBs) and electron-deficient multicenter bonds (EDMBs), with special emphasis on molecules and solids primarily composed of electron-rich elements. We show that both types of multicenter bonds have the same origin, but a different mechanism of formation upon increasing electronic density; e.g. upon reduction, increase of pressure, or chemical substitution of elements by their heavier analogs. In addition to our recent demonstration of the formation of E MDBs in electron-rich systems, such as pnictogens, chalcogens, and chalcogenides related to phase change materials (see H. H. Osman *et al.*, *J. Mater. Chem. C*, 2024, **12**, 10447–10474), here we present other examples of solids with electron-rich elements forming E MDBs and ERMBs. We conclude that E MDBs can occur not only as three-center–two-electron (3c–2e) bonds in molecules but also as linear or quasi-linear two-center–one-electron (2c–1e) bonds in extended solids. In addition, we propose that pure ERMBs can only occur as linear or quasi-linear three-center–four-electron (3c–4e) bonds. All these claims suppose a change of paradigms regarding the current understanding of ERMBs and E MDBs. To understand the formation of ERMBs and E MDBs in electron-rich elements, we show some of the simplest geometries of linear E MDBs and ERMBs along one (1D), two (2D), and three (3D) dimensions that can be found in the hypercoordinated multicenter units of molecules and solids with electron-rich elements and propose a new way of notation of these hypercoordinated units. Finally, we show that both types of multicenter bonds, in general, do not violate the doublet/octet rule.

Received 17th October 2024,
Accepted 28th January 2025

DOI: 10.1039/d4tc04441j

rsc.li/materials-c

1. Introduction

The nature of the chemical bonding in the crystalline phases of phase change materials (PCMs) has been debated for over 70 years due to their extraordinary properties. In recent works,^{1,2} it has been proposed that the octahedrally-coordinated α -Po and β -Po phases observed in pnictogens and chalcogens, as well as the crystalline phases of PCMs, feature the old electron-deficient multicent bond (EDMB). This bond is the generalized

version of the three-center–two-electron (3c–2e) bond proposed by Rundle and Longuet-Higgins in the 1950s^{3–6} and it is known to be typically present in materials with electron-deficient elements (hydrogen and elements of groups 1, 2, and 13), such as the hydrogenonium ion H_3^+ , elemental boron, boranes, and BaAl_4 among others.^{7–10} Noteworthy, W. Lipscomb was awarded the Nobel Prize in Chemistry in 1976 for the explanation of boranes with the use of 3c–2e bonds.¹¹

The electron-deficient character of the bonds in the octahedrally-coordinated α -Po and β -Po phases of group-15 and -16 elements has been evidenced by the small value of the charge density at the bond critical point and the electron localization function (ELF) between two bonded atoms as well as by the number of electrons shared between two atoms (ES) obtained from quantum mechanical calculations.^{1,2} This last magnitude is calculated as two times the delocalization index (DI) obtained from the quantum theory of atoms in molecules (QTAIM) as already done in previous works.^{12–15} It has also been shown that the ES values combined with the renormalized number of electrons transferred between two atoms (ET),

^a Instituto de Diseño para la Fabricación y Producción Automatizada, MALTA Consolider Team, Universitat Politècnica de València, 46022 València, Spain.
E-mail: fjmanjon@fis.upv.es

^b Instituto de Ciencia de los Materiales de la Universitat de València, MALTA Consolider Team, Universitat de València, 46100 Valencia, Spain.
E-mail: hussien.helmy@uv.es

^c Chemistry Department, Faculty of Science, Helwan University, 11795 Cairo, Egypt
^d Departamento de Física, MALTA Consolider Team, Universidad de La Laguna, 38205 La Laguna, Tenerife, Spain

† Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d4tc04441j>

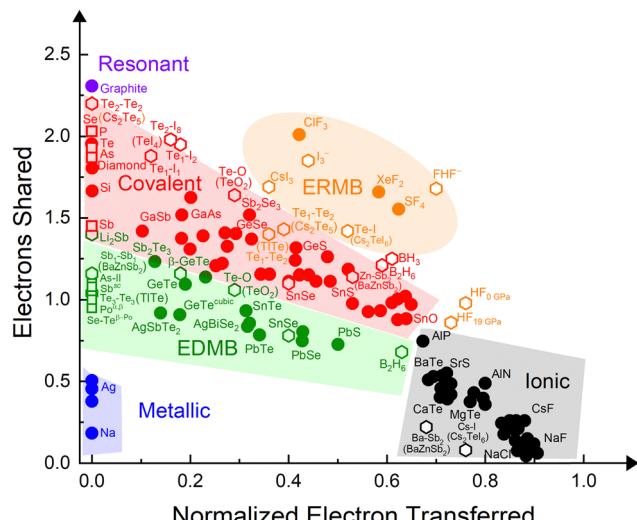


Fig. 1 2D map with the number of electrons shared (ES) vs. the normalized number of electrons transferred (ET) showing the chemical bond classification in materials. The red, black, and blue regions correspond to the classical covalent, ionic, and metallic bonds, respectively. In addition, the map shows the orange and green regions of materials with electron-rich multicenter bonds (ERMBs) and electron-deficient multicenter bonds (EDMBs), respectively. Circles correspond to materials studied in previous studies.¹⁴ Pnictogens and chalcogens studied in a previous work at different pressures² are shown as open squares. As observed, the structures of pnictogens and chalcogens with octahedral coordination are located in the green region of EMBs as well as PCMs, such as GeTe, SnTe, and PbS.

typically calculated as the Bader charge of the cation divided by the nominal oxidation state of the cation, can be used in a two-dimensional (2D) ES vs. ET map (Fig. 1) to identify the different types of bonds in materials.^{12–15} Importantly, the density-based ES vs. ET map has been demonstrated to be equivalent to an orbital-based map, where the ES value is calculated as two times the integrated crystal orbital bond index (ICOBI), and the ET value is calculated as the normalized Löwdin charge.^{15,16} In this context, we have also provided the orbital-based ES vs. ET map in Fig. S1 in the ESI[†] to complement the density-based map. Although both maps provide the same information, some examples are located better on one map than on the other. Consequently, we believe that the ES vs. ET map, or any other equivalent map, could help to locate the different kinds of bonds in different regions of the map. In particular, we are going to show that it can be used to distinguish between the two known types of multicenter bonds, the EMB and the electron-rich multicenter bond (ERMB) from the classical covalent, ionic, and metallic bonds.

The ERMB, also known as the hypervalent bond,¹⁷ is the generalized version of the three-center-four electron (3c–4e) bond proposed by Rundle and Pimentel in the 1950s.^{18,19} This bond has been assumed to be typically present in materials with electron-rich elements (elements of groups 15 to 18), such as hypervalent molecules I_3^- , XeF_2 , XeF_4 , XeF_6 , simple cubic (sc) Sb, and several Zintl phases, such as Li_2Sb , $BaZnSb_2$, $TlTe$.²⁰ It must be mentioned that the very existence of ERMBs has been questioned in several works.^{21,22}

According to the ES and ET values in Fig. 1, covalent bonds with a bond order of 1 (red region in Fig. 1) extend from pure covalent bonds (ET = 0 and ES ≈ 2) to ionic-covalent bonds or polar covalent bonds (ET ≈ 0.65 and ES ≈ 0.9).^{2,14} In contrast, EMBs in the octahedrally-coordinated α -Po and β -Po phases of pnictogens and chalcogens (also with ET = 0) as well as PCMs (ET ≠ 0), such as β -GeTe, SnTe, and PbS, extend along the same range of ET values, but are characterized by much smaller values of ES (ES ≈ 1) (green region in Fig. 1 and Fig. S1, ESI[†]). Note that ERMBs are characterized by having a bond order of 0.5 since they are longer than covalent bonds. We will show later that a similar situation is found for a molecule with ERMBs, such as diborane (B_2H_6),²³ whose central B–H bonds are also located in the green region of Fig. 1 and Fig. S1 (ESI[†]).

Similar to EMBs, ERMBs are longer than covalent bonds and were supposed to have a bond order of 0.5 according to the Pimentel model;^{18,19} however, ERMBs are typically shorter than EMBs and feature a bond order higher than those of EMBs, as has been recently shown for polyiodides.²⁴ However, the ES values of ERMBs are larger than those of ionic-covalent bonds of similar polarity (similar value of ET). This result allows us to locate ERMBs in the orange region of Fig. 1 and Fig. S1 (ESI[†]), as we will show later for molecules with ERMBs, such as $[FHF]^-$ and XeF_2 , and for solids, such as CsI_3 , whose H–F, Xe–F, and I–I bonds, respectively, are in good agreement with recent findings in hypervalent molecules.¹⁵

The EMB model for PCMs proposed in ref. 1 and 2 is in contrast to the two previous bond models for PCMs: the hypervalent and metavalent bond models. The hypervalent bond model considers that the crystalline phases of PCMs, which feature hypercoordinated electron-rich elements, are characterized by the ERMBs; however, this does not agree with the position of bonds in PCMs in the ES vs. ET map. On the other hand, the metavalent bond model considers that the crystalline phases of PCMs are characterized by a new bond type, the metavalent bond, which is a two-center-one-electron (2c–1e); *i.e.*, an electron-deficient bond, whose position in the ES vs. ET map is similar to that of the position of the EMBs. Noteworthy, the supporters of the metavalent bond model for PCMs do not consider the multicenter character of the bond in PCMs although they do not discard that it can have a multicenter character.^{13–15,25,26} In recent works, we have stressed that if the multicenter character is added to the electron-deficient nature of the metavalent bond, then the metavalent bond becomes naturally the EMB.^{1,2,27}

It must be clarified that the proponents of the metavalent bond model have not clarified the position of EMBs in the 2D ES vs. ET map. Instead, the authors of ref. 1 and 2 claim that the new metavalent bond is nothing else but the old EMB. Therefore, the properties of the EMB are those that have been attributed to the new metavalent bond: (i) hypercoordination (violation of the 8 – N rule); in other words, a higher atomic coordination than that expected for compounds with covalent $p\sigma$ -bonds that obey the 8 – N rule, (ii) relatively low band gaps and shiny metallic luster, (iii) moderate electrical conductivity, (iv) extremely high optical dielectric constants and Born effective

charges, (v) low-frequency optical phonons and high Grüneisen parameters, (vi) low thermal conductivity, (vii) brittleness (due to the directional character of EDMBs), and (viii) high probability of multiple emission events in laser-assisted field evaporation measurements (due to the softer character of EDMBs than that of the covalent ones).[‡]

In ref. 1 and 2, it has been also shown that the octahedrally coordinated α -Po and β -Po phases of pnictogens (As, Sb, Bi) and chalcogens (Se, Te) at high pressure (HP) feature linear EDMBs formed by a mechanism that transforms the primary covalent bond and the secondary weak non-covalent bond (present in the low-coordinated phases of these elements at room pressure (RP)) into EDMBs at HP. It was commented that the mechanism of EMB formation is different from that of ERMB formation; however, these two mechanisms were not discussed in considerable detail in those works.

The present work is the second part of the manuscript already sent for publication on December 2023,¹ which we have decided to split into two parts (ref. 2 and this work) for the sake of clarity. This work presents a unified multicenter bond theory which is aimed at discussing at length the similitudes and differences between ERMBs and EDMBs, mainly in electron-rich elements; the common origin and different formation mechanisms of both multicenter bonds; the geometries of the hypercoordinated units of electron-rich elements with both types of multicenter bonds; and the compliance of the doublet/octet rule, in general, in the atoms participating in both ERMBs and EDMBs. To facilitate the understanding and differentiation of both kinds of multicenter bonds in electron-rich elements, we provide and discuss the crystalline structures of different molecules and solids with different bond types (even a solid with a mixture of both EDMBs and ERMBs in a single element, *i.e.*, homoatomic bonds, within the same crystalline structure). The clear distinction between ERMBs and EDMBs will help us to explain the structures and properties found in many important technological materials, including PCMs, thermoelectrical and photovoltaic materials, Zintl phases, intermetallics, polyhalides, and cluster compounds at RP as well as the structures and properties found in several polyhalogen anions and the atomic/polymeric phases of nitrogen and hydrogen at HP.

To achieve this, we have performed first-principles simulations based on density functional theory (DFT) for several materials (further details are provided in the ESI[†]). In particular, a comparison of the experimental and simulated structural data of the different materials studied in this work is provided in Table S1 of the ESI.[†] The theoretical bonding descriptors between two atoms are provided for density- (Table S2 of the ESI[†]) and orbital-based methods (Table S3 of the ESI[†]). In particular, the density-based method using QTAIM¹² allows

[‡] The properties of a solid should not be confused with those of its bonds. Solid properties emerge from the collective behavior of bonds in the structure. In solids with a single bond type, like rocksalt-type PCMs, solid properties may reflect bond properties. However, in solids with multiple bond types, such as B_2H_6 , the solid's properties result from the combined response of all bonds, making it difficult to directly infer individual bond properties.

us to obtain the Bader atomic charges as well as ES and ET values using the CRITIC2 program,²⁸ while the orbital-based method allows us to obtain the Löwdin charges and ES and ET values using the LOBSTER program.²⁹ Hereon, the different values of bonding descriptors in different materials will be noted as X.X [X.X] as obtained from CRITIC2 [LOBSTER] program and these values can be found in Table S2 [Table S3, ESI[†]]. Note that the ES and ET values obtained from CRITIC2 and LOBSTER are different; however, in general, they provide similar information if we compare Fig. 1 and Fig. S1 (ESI[†]), as already shown in previous works.^{2,16,25} Therefore, it can be concluded that our results on ES and ET values obtained from DFT simulations are robust since they do not depend on density- or orbital-based methods as long as DFT simulations are considered to be valid for the description of the studied systems.

Here we want also to notice that the distinction of the different types of multicenter bonds (ERMBs and EDMBs) from ionic-covalent bonds is mainly done by taking into account the different bond lengths of multicenter and ionic-covalent bonds and with the use of ES and ET values. Note that multicenter bond indices for solids have not been well developed yet. In this context, it must be mentioned that despite multicenter bond indices have been calculated for molecules since 1990,^{22,30–32} and extended to aromatic molecules,^{33,34} multicenter bond indices for solids have been barely explored since the results are not so clear with correlated wavefunctions³⁵ and are not available in the CRITIC2 code yet. A multicenter bond index for solids has been proposed in LOBSTER, the integrated crystal orbital bond index for three centers, ICOBI(3c);³⁶ however, its validation for solids, like PCMs is doubtful,³⁷ as already discussed in ref. 2.

2. Origin and mechanism of formation of ERMBs and EDMBs

Multicenter bonds, either electron-deficient or electron-rich, are bonds extended beyond two centers. Therefore, the simplest multicenter bond is necessarily the three-center bond as found in trimers. Consequently, the origin of multicenter bonds in trimers, either electron-rich, like in I_3^- , or electron-deficient, like in H_3^+ , must come from the interaction of primary bonds, like intramolecular bonds in I_2 and H_2 molecules, and secondary non-covalent bonds, like intermolecular bonds between I_2 and I^- and between H_2 and H^+ . The formation of three-center bonds requires a *trans* influence between both the primary and secondary bonds; that is, a perturbation of the primary bond (typically resulting in lengthening and weakening) as the secondary bond shortens and strengthens. In the end, the original primary and secondary bonds (both are two-center interactions) tend to be equal in length and strength (provided that they are formed by the same elements, such as in I_3^- and H_3^+) to give rise to the multicenter (three-center) bond in the trimer.

Examples of the formation of 3c–4e bonds in trimers coming from the interaction between primary and secondary bonds



were already provided in 1998 by Landrum and Hoffmann, who realized that the formation of linear ERMBs in trimers with central Sb and Te atoms was promoted when the primary bond length of the trimer was above a certain value.³⁸ These data were further analyzed by Lubchenko and coworkers, who stated more clearly that the formation of ERMBs was promoted with the decrease of the trimer mean bond length as the mass density increases.³⁹ The same tendency to form ERMBs as the trimer total length decreases was found in the formation of ERMBs in I_3^- trimers in different compounds by Savastano,⁴⁰ who revised the old work of Money-Slater.⁴¹ In all these papers, it was shown that both primary and secondary bonds tend to equalize when the ERMB is formed. In addition, Papoian and Hoffmann suggested that ERMB formation could also be induced by reduction, *i.e.*, providing electrons to the system.²⁰ Altogether, these data suggest that the formation of ERMBs by equalization of original short primary and long secondary bonds is promoted when the electronic density increases.

On the other hand, it has been recently concluded that EMB formation in pnictogens and chalcogens can be also induced by pressure and by atomic substitution by heavier elements in a group;² *i.e.*, two mechanisms that also lead to an increase in electronic density. It must be stressed that a similar conclusion was provided by Grochala *et al.*, who considered that pressure promoted the formation of multicenter bonds.⁴² Consequently, we can tentatively conclude that, in general, the formation of multicenter bonds is promoted by any means that leads to an increase in the electronic density (either by pressure, substitution by heavier elements, or reduction). Conversely, it can be also concluded that multicenter bonds split into a primary and a secondary bond when the electronic density goes below a given value (induced either by decreasing pressure, substitution by lighter elements, or oxidation); a sentence that is in line with the results of Lubchenko and coworkers.³⁹

To illustrate the general simplest process of multicenter bond formation on increasing electronic density, we present a scheme of how multicenter bonds can be formed from initial primary bonds and secondary interactions/bondings *via* a gradual process with three stages (see Fig. 2). The evolution of the distances of a primary strong bond (d_1) and a relatively weak secondary interaction (d_2) is shown in Fig. 2a, while the evolution of the optical phonon frequencies of a material that develops a multicenter bond with increase of the electronic density is shown in Fig. 2b.

Let us first comment on the evolution of bond lengths with increasing electronic density during the formation process of multicenter bonds (Fig. 2a). In the first stage of multicenter bond formation, the ratio between the distances of the secondary interaction and the primary bond is so large (typically $d_2/d_1 > 2$) that both bonds/interactions are not linked; *i.e.*, the strong primary bond (typically of ionic-covalent nature) is not affected by the presence of the weak secondary interaction (typically being a non-covalent interaction of electrostatic nature). In the second stage, the ratio between the distances of the secondary and primary bonds are comparable (typically $1 < d_2/d_1 < 2$) due to the considerable shortening of the secondary bond length (now the secondary interaction can be considered

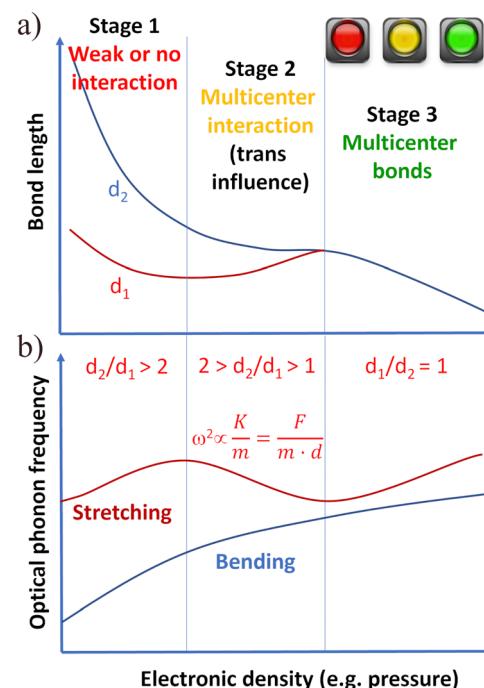


Fig. 2 General behavior of the (a) bond lengths and (b) optical phonon frequencies of a material with a primary and a secondary bond, which are linked by a *trans* influence, and therefore lead to the multicenter bond formation upon the increase of the electronic density; e.g. upon reduction, increase of pressure, or chemical substitution of an element by a heavier analog.

a bond) as the electronic density increases. In this stage, the secondary bond perturbs (*trans* influence) the primary bond, so both bonds tend to a similar (sometimes equal) value of bond length as the electronic density increases. The *trans* influence is a well-known phenomenon in Chemistry[§] by which the primary

[§] Notice that it would be more appropriate to say that the primary and secondary bonds mutually influence each other during the formation of a multicenter bond. This is a consequence of Newton's third law which states that for every action there is an equal and opposite reaction. Therefore, in any interaction between two objects A and B, the "action" and "reaction" refer to forces, so if object A exerts a force on object B, then object B exerts an equal amount of force on object A in the opposite direction. However, the weak secondary bond evolves as expected during the formation of a multicenter bond (it increases in strength and decreases in bond length when the multicenter interaction starts in stage 2), while the strong primary bond evolves in an unexpected or anomalous way during the formation of a multicenter bond since it decreases in strength and increases in length in stage 2. Therefore, in this manuscript, we prefer to say that the secondary bond influences the primary bond as has been understood for decades in the chemical bond terminology of supramolecular interactions. We hope that maintaining this terminology would lead chemists to a better understanding of the process of formation of multicenter bonds despite being true that in any interaction there is a mutual influence as expected from Newton's third law. Notice that our point of view is similar to the one we commonly use for the fall of an object in a gravitational field. For instance, we say that a small object falls to the Earth and not that the object and the Earth fall towards the center of mass of the compound system (object + Earth). As usual, if the change of the center of gravity of the object is much larger than the change of the center of gravity of the Earth in their travel towards the center of mass of the compound system, we use to say that the object falls to the Earth (although we all implicitly know that both object and Earth fall or attract to each other).



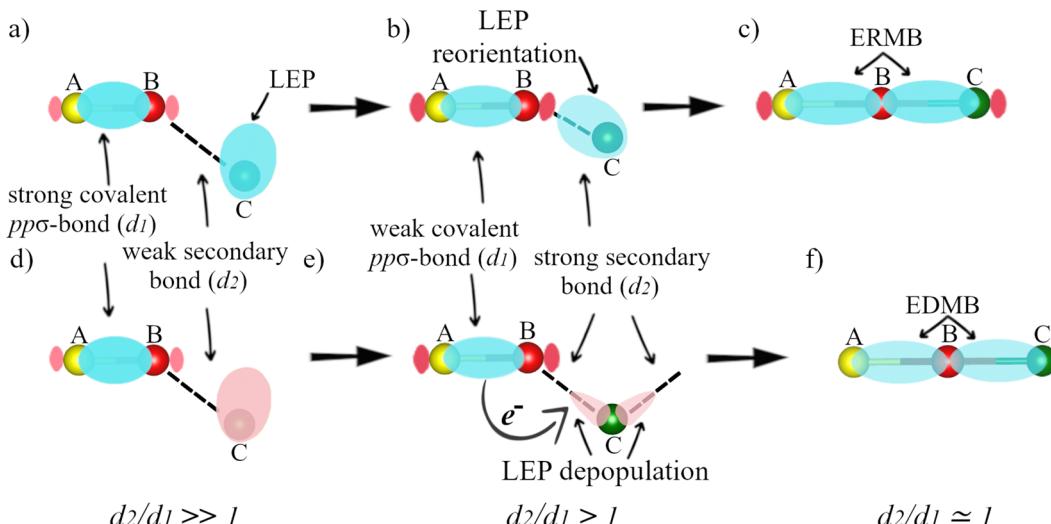


Fig. 3 (a)–(c) Schematics of the mechanism of a 3c–4e bond formation in a system with three electron-rich atoms (A, B, and C) as the electronic density increases. (d)–(f) Schematics of the mechanism of a 3c–2e bond formation in a system with three electron-rich atoms (A, B, and C) as the electronic density increases. At low electronic densities (a) and (d), A and B atoms are assumed to be linked by a primary iono-covalent bond and B and C atoms are linked by a secondary non-covalent interaction/bond in which a LEP is involved.

bond length typically increases at the expense of the decrease of secondary bond length and it has been documented in a number of papers during the formation of ERMBs^{24–26,38} and E MDBs.^{1,2} In the third stage, beyond a critical electronic density, both bonds become equal or almost equal in bond length (typically $d_2/d_1 < 1.05$) and strength and behave similarly upon an increase in electronic density. When this behavior occurs, a multicenter bond has been formed.

It must be stressed that the presence of three stages during the formation of E MDBs was clearly shown in ref. 1 and 2 and that the same number of stages was noted during the formation of 3c–4e bonds in several systems, *e.g.* in $[\text{HF}]^-$ or HF_2^- ,⁴³ and during the breaking of 3c–4e bonds; *e.g.* the C–C≡N of acetonitrile during a $\text{S}_{\text{N}}2$ reaction.⁴⁴ The observation of these three stages in the formation of multicenter bonds contrasts with the results of several works that suggested that it would be difficult to distinguish between the scenario with primary and secondary bonds and the scenario with multicenter bonds.^{38,39,45}

The three stages of the multicenter bond formation process are also reflected in the evolution of the vibrational modes of the materials with increasing electronic density (see Fig. 2b). In the first stage, the increase in electronic density leads to a normal increase in the optical phonon frequencies of both high-frequency stretching phonons and low-frequency bending phonons (the last ones derived from the acoustic phonons of the Brillouin zone edge due to folding of the Brillouin zone). In the second stage, the *trans* influence leads to an anomalous softening of the stretching phonons. This softening can be understood if we consider that the square of the frequency of stretching phonons is proportional to the bond force constant per unit length, K , and inversely proportional to the mass, m , of the linked atoms or equivalently proportional to the bond force constant, F , and inversely proportional m , and bond length, d ,

of the linked atoms (see equation in Fig. 2). Therefore, the anomalous increase in the primary bond length (in addition to the loss of charge of the primary bond in the mechanism of E MDB formation (see Fig. 3d–f) explains the softening of the stretching phonons related to the primary bond. Finally, in the third stage, there is a re-hardening of the stretching phonons due to the normal behavior of the bond distance and charge density of the multicenter bonds.^{1,2}

We want to stress that the different behavior of the frequency of acoustic and bending phonons compared to that of stretching phonons as a function of pressure (*i.e.*, one process that allows increasing the electronic density) during the whole process of formation of E MDBs has been already discussed regarding the bonding present in PCMs.^{1,2,46} We also want to note that the behavior described in Fig. 2b is a general one and does not consider phonon anharmonic interactions or electron–phonon interactions that can make each system evolve in a particular way. It is noteworthy that the three-stage process in the multicenter bond formation is consistent with the distribution of the electronic charge of atoms in three spheres: the core, valence, and van der Waals spheres, and the penetration indices of these spheres as discussed by Echeverría and Álvarez.^{47,48} In particular, the three-stage formation process of multicenter bonds with the penetration of the valence spheres is commented on Fig. S2 of the ESI† concerning the formation of the I_3^- anion from the original I_2 and I^- entities, which is discussed in the next section.

All in all, the aim of Fig. 2 is to show that both ERMBs and E MDBs have a common origin; *i.e.*, the existence of a primary bond and a secondary interaction that are linked by the *trans* influence, and also a similar formation process. Most specifically, Fig. 2 shows the simplest way of forming ERMBs and E MDBs; *i.e.*, *via* a second-order (or even higher-order) phase transition in



which there is a progressive and gradual transformation of the primary bond and the secondary interaction into multicenter bonds *via* an intermediate stage characterized by the occurrence of secondary bonds and a *trans* influence process that involves a charge reorganization in both bonds. Note that the above description does not preclude the possible formation of ERMBs and EDMBs upon increase of electronic density *via* a reconstructive (first-order) phase transition in which stage 2 can be bypassed.

At this point, it is important to stress that the mechanisms of ERMB and EMB formation are different.^{1,2} The mechanism of ERMB formation has been schematized in ref. 45 and 49, while the mechanism of EMB formation has been schematized in ref. 1 and 2. Fig. 3 schematically shows the two mechanisms of multicenter bond formation for three electron-rich atoms (A, B, and C) as the electronic density increases. Since we only use three atoms for simplicity, Fig. 3a–c illustrates the formation of an ERMB (3c–4e bond), while Fig. 3d–f illustrates the formation of an EMB (3c–2e bond). We make this comment because we will later address the case of 2c–1e EDMBs extended to more than three centers (a case that we will show cannot occur in ERMBs), so Fig. 3d–f is a simplified version of the real situation in extended systems with a 2c–1e ERMBs.

Let us start with the ERMB formation mechanism

The formation of a 3c–4e bond was addressed by Lee and Elliott,^{45,49} who proposed that the two non-bonding electrons of the stereochemically active LEP of atom C (Fig. 3a) become more and more aligned with the A–B bond as the electronic density increases; *i.e.*, there is a LEP reorientation along the B–C direction (Fig. 3b). This process goes on until the three atoms (A, B, and C) are fully (or almost fully) aligned once the A–B–C 3c–4e bond is formed (Fig. 3c). In this process, the non-bonding electrons of the LEP in Fig. 3a transform in a progressive way in bonding electrons in Fig. 3c as the electronic density increases. The secondary non-covalent interaction that leads to the ERMB formation is traditionally considered to be caused by the interaction of the LEP (donor or Lewis base) of atom C and the anti-bonding orbital (σ^* , acceptor or Lewis acid) associated with the primary ionocovalent $\text{pp}\sigma$ -bond between A and B atoms. Typically, there are two models of this donor–acceptor secondary interaction,⁵⁰ the charge transfer (LEP– σ^*) model, also noted as $\text{n} \rightarrow \sigma^*$ model,⁵¹ and the σ -hole bond model.^{52,53} The LEP– σ^* bond model considers that the LEP charge is transferred to σ^* as the electronic density increases, while the σ -hole bond model considers that there is no charge transfer from the LEP to σ^* and that there is only an electrostatic interaction between an electronegative (nucleophilic) region, the LEP, and the electropositive (electrophilic) region, known as the σ -hole, situated at the end of the ionocovalent bond. This last model considers that there is no charge transfer but only a charge shift and polarization. In any case, both models account for the formation of the ERMBs and it is commonly believed that the σ -hole bond model represents best the initial stages of ERMB formation while the LEP– σ^* bond model represents the final stages of ERMB formation. We must also note that, in the last years, the formation of ERMBs in

hypercoordinated structures has also been understood as the delocalization of electrons of LEPs into bonding or anti-bonding orbitals according to the “increased-valence” theory.⁵⁴ Therefore, it is clear, as we propose, that ERMB formation involves the transformation of non-bonding electrons in LEPs into bonding electrons in the ERMB as shown in Fig. 3a–c.

Let us continue now with the EMB formation mechanism

As suggested by Manjón and coworkers for the formation of 3c–2e (also applied to extended 2c–1e bonds),^{1,2} the non-bonding electrons of the stereoactive LEP of atom C (Fig. 3d) in general do not participate in bonding as the electronic density increases. In this scenario, the LEP becomes distorted, depopulated, and delocalized as the electronic density increases (despite a part of the LEP is reoriented along the B–C direction of the secondary bond, Fig. 3e). In this way, the LEP (if there exists as in electron-rich elements) progressively loses its stereoactivity until the three atoms (A, B, and C) are fully (or almost fully) aligned once the A–B–C 3c–2e bond is formed (Fig. 3f). Along this process, the non-bonding electrons of the LEP, in general, remain non-bonded and the charge needed to form the new B–C bond (of the A–B–C 3c–2e bond) comes from the primary ionocovalent A–B bond (see curved arrow in Fig. 3e). In other words, there is a net charge transfer from the ionocovalent A–B bond to the newly formed B–C bond as the electronic density increases until the charges of both A–B and B–C bonds equalize and the multicenter A–B–C 3c–2e bond is formed.

It must be stressed that the secondary interactions that lead to the EMB formation cannot be explained either with the LEP– σ^* bond model or with the σ -hole bond model since these two models do not consider a charge transfer from the primary bond to the secondary interaction. On the other hand, it must be mentioned that the mechanism of EMB formation might be explained with the “increased-valence” theory,⁵⁴ which accounts for the electron delocalization of non-bonding electrons from LEPs, provided that the theory is extended to include the delocalization of bonding electrons from bonding electron pairs (BEPs), *i.e.*, those present in ionocovalent bonds.

We want to note that the mechanism we have proposed for the formation of EDMBs in electron-rich elements in Fig. 3d–f, in which LEPs do not participate, in general, in the formation of the secondary and multicenter bond, is consistent with the formation of EDMBs in electron-deficient elements since electron-deficient elements do not have LEPs that can participate in secondary and multicenter bonds. Therefore, the LEPs play a major role in the formation of secondary and multicenter bonds in ERMBs (see Fig. 3a–c), while LEPs play a minor role in the formation of secondary and multicenter bonds in EDMBs (see Fig. 3d–f).

We want to close this general description of multicenter bond origin and formation by stating that the two different mechanisms of ERMB and EMB formation are reflected in the different evolution of the ES and ET values for the primary and secondary bonds/interactions as the electronic density increases (see Fig. S3 of the ESI[†]). During the formation of



both ERMBs and EDMBs, the ES and ET values tend to equalize in both cases; however, their final values are completely different in ERMBs and EDMBs. On the one hand, ES values should be between 1.4 and 1.9 in ERMBs while they are typically around 1.0 in EDMBs. On the other hand, ET values can be considered to increase during the ERMB formation and be more or less constant or have a small change during the EDMB formation.

3. Examples of the mechanisms of ERMB and EMDB formation

In this section, we show examples that illustrate the mechanisms of ERMB and EMDB formation. The two mechanisms can be illustrated by the formation of two well-known entities, the 3c–4e bonds in I_3^- polyanion and the 3c–2e bonds in diborane (B_2H_6) molecule, as schematized in Fig. 4a and b, respectively. Note that we have used in those figures the formalism of the σ -hole bond model in both cases to account for the electrostatic interaction that is present at the initial stage (stage 1 in Fig. 2) of ERMB and EMDB formation in both systems, despite the σ -hole bond model cannot account for the charge transfer occurring in stage 2 of the EMDB formation, as previously discussed.

I_3^- polyanion

The ERMB formation in the I_3^- polyanion can be considered as the result of the approach of an I^- ion to an I_2 molecule as the electronic density increases (Fig. 4a).^{40,50,55} As they approach, a *trans* influence of the secondary bond between the I_2 molecule and the I^- ion into the covalent bond of the I_2 molecule occurs. As already commented, the ERMB formation in I_3^- can be understood in the light of the three current bond models. According to the donor–acceptor charge-transfer model, a charge transfer from the donor unit (I^-) to the acceptor unit

(the anti-bonding orbital of the covalent bond in I_2) occurs as they approach each other. In other words, the anti-bonding orbital of the covalent bond of the I_2 molecule is populated due to the charge transfer from the LEP of the I^- ion. This *trans* influence occurs until both the intramolecular I–I distance of the I_2 molecule and the intermolecular I–I distance between the I^- ion and one of the I atoms in the I_2 molecule become almost equal.^{40,56} Alternatively, the mechanism of ERMB formation can be understood in the light of the σ -hole bond model of secondary bonds. In this model, the negatively charged I^- ion (nucleophile) is attracted to the electropositive (electrophile) region known as the σ -hole, which is situated at the end of the I–I covalent bond within the I_2 molecule. Contrary to the charge-transfer model, when the electrophilic and nucleophilic entities approach each other, there is no outright transfer of charge from the donor to the acceptor unit. Instead, a polarization of charges occurs within the I_2 molecule due to the *trans* influence caused by the proximity of the I^- ion.

Specifically, as the I^- ion approaches the closer I atom of the I_2 molecule (now it becomes the central I atom of the I_3^- polyanion), the I^- ion shares part of its charge with the central I atom. This sharing induces a shift of part of the charge of the central I atom towards the other I atom of the I_2 molecule, which is positioned opposite to the I^- ion, so a symmetric ($- + -$) charge configuration is observed in the 3c–4e bond. This mechanism of ERMB formation explains the more electro-negative external parts of this multicenter bond, as originally proposed by Rundle and Pimentel,^{18,19} and also the large values of the normalized number of electrons transferred (ET) of molecules with ERMBs, such as XeF_2 , CF_3 , and SF_4 .²⁵ In fact, the values of ES (above 1.6) and ET (0.45, obtained as the difference in Bader charges between the central and external atoms, see all details in Table S2, ESI†) for the I–I bond in the I_3^- ion (see Fig. 4a) allow us to locate this bond in the orange region of Fig. 1 corresponding to the ERMB, like XeF_2 . The same description is found when the values of the bonding

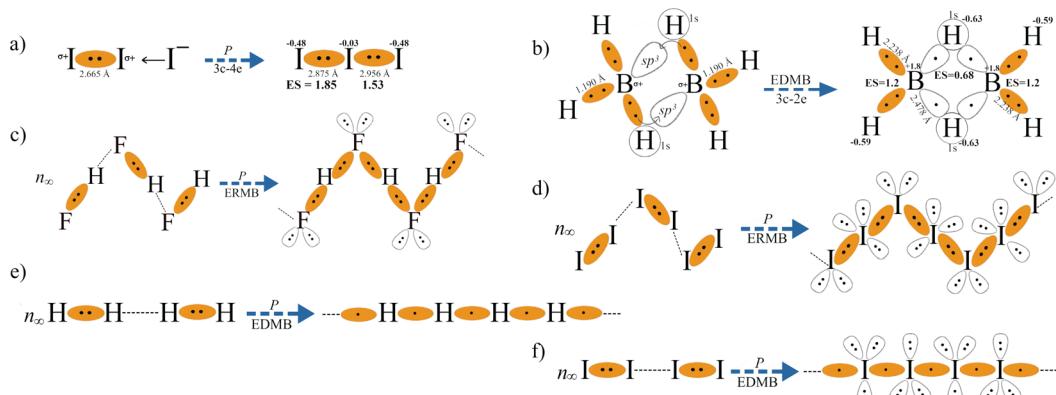


Fig. 4 Scheme of the formation mechanism of (a) I_3^- polyanion and (b) B_2H_6 molecule using the formalism of the σ -hole bond model for secondary bonds. The charge transfer from the primary covalent bond towards the secondary bond in (b) is plotted with an arrow. (c) and (d) Schemes of the pressure-induced polymerization of HF and I_2 molecules which leads to an infinite zigzag chain with quasi-ERMBs. (e) and (f) Schemes of the pressure-induced polymerization of H and I atoms which leads to an infinite linear chain with EDMBs. Here we show some details on bond distances, Bader atomic charges, ET, and ES of various bonds in I_3^- and B_2H_6 that can be found in Table S2 (ESI†). LEPs are depicted on the right side of (d) and (f) to illustrate the total electron distribution in the infinite chains with multicenter bonds in electron-rich elements.



descriptors are obtained from orbital-based methods (see Fig. S1 and Table S3, ESI[†]). Therefore, there is an equivalence of both ES vs. ET maps obtained from density-based (Fig. 1) and orbital-based (Fig. S1, ESI[†]) methodologies. Finally, the third model suggests that the formation of hypercoordinated structures with ERMBs can be understood within the “increased-valence” theory as caused by the delocalization of electrons of LEPs into bonding or anti-bonding orbitals.⁵⁴ In this context, we think that the σ -hole bond model and the “increased-valence” models seem to be more appropriate than the charge-transfer model to describe the ERMB formation. Therefore, we conclude that the ERMB formation does not involve a charge transfer between the primary and secondary bonds, unlike in EDMBs.² In other words, the primary bond in the ERMB does not lose charge, thus resulting in a rather large ES value compared to EDMBs, as found for I_3^- and XeF_2 (see ref. 25 and Tables S2 and S3, ESI[†]). It is the delocalization of non-bonding electrons at LEPs into the secondary bond which provides the charge for the formation of the ERMB as shown in Fig. 3a-c.

In connection with the formation of ERMBs in I_3^- it may be worth pointing out that it has been suggested that the charge shift mechanism during the ERMB formation in trimers that leads to the symmetric $(- + -)$ charge configuration in the 3c–4e bond avoids the central atom of the three-center entity to severely violate the octet rule.^{52,56–59} Indeed, the shift of charge to the external parts of a hypervalent trimer, such as I_3^- , results in an 8 + 2 charge of the central atom (see Fig. 5). This means that, excluding the first electronic sphere of core electrons (dark blue sphere in Fig. 5), in a first approximation we can consider that the central atom avoids violating the octet rule since the second electronic sphere (light blue in Fig. 5),

corresponding to the valence electrons, only contains 8 electrons. Only when the third electronic sphere (light red in Fig. 5) is included, there is an apparent violation of the octet rule since this sphere contains 2 electrons.

It is important to note that Echeverría and Álvarez have recently justified the distribution of the electronic charge in different spheres as mentioned in the previous paragraph. These authors have ordered the different spheres in order of decreasing electron density and named them core, valence, and van der Waals spheres.^{47,48} They have used the different spheres to introduce the concept of penetration index to understand the chemical bonding between the strong ionic-covalent bonds and the weak van der Waals interactions, including hydrogen bonds and secondary non-covalent interactions present in materials that lead to the formation of ERMBs and EDMBs, such as those of HF_2^- and B_2H_6 , respectively. In this context, the large bond length of ERMBs formed by the two largest spheres depicted in Fig. 5 can be understood as the sum of the valence sphere (typical of the ionic-covalent bonds^{47,48}) and an additional small van der Waals sphere (typically with a radius 0.2–0.4 Å larger than the valence sphere). In such a way, the central atom of an ERMB would always have 8 electrons in the valence sphere, thus satisfying the octet rule, and the additional charge would be located at the additional van der Waals sphere, thus avoiding the violation of the octet rule.

A similar procedure allows us to explain the linear ERMB formation and charge distribution in other 1D linear molecules, such as XeF_2 and HF_2^- , in 2D molecules, such as XeF_4 , and even in 3D molecules, such as SF_6 and XeF_6 . We propose that the electronic configuration of the central atom in XeF_4 and XeF_6 can be thought to be 8 + 4 and 8 + 6, respectively.⁶⁰ Therefore, Fig. 5 can be extended in 2D and 3D to explain how electrons are distributed in the different electronic spheres in XeF_4 and XeF_6 molecules so that the central Xe atom avoids the violation of the octet rule, because only 8 electrons are in the valence sphere while the rest are occupying the outer van der Waals sphere.

Once introduced the concepts of the core, valence, and van der Waals spheres,^{47,48} the three-stage mechanism of ERMB formation in the I_3^- anion can be understood with the use of the valence sphere, as described in Fig. S2 (ESI[†]). It can be observed that the multicenter interaction (stage 2) starts when $d_2/d_1 = 2$ in the formation of ERMBs, such as in the I_3^- anion and the $[FHF]^-$ molecule, as discussed in the ESI.[†]

B_2H_6 molecule

Now that the formation of ERMBs has been clarified, we will deal with the EDMB formation in the B_2H_6 molecule (Fig. 4b), which can be considered as a two-proton attack on the B–B bond of the $B_2H_4^{2-}$ unit,⁶¹ but it can also be seen as two borane (BH_3) molecules with three covalent B–H bonds that interact together *via* secondary interactions in a dimerization process.^{62,63} According to the σ -hole bond model of secondary interactions, a σ -hole appears at the end of each atom involved in the covalent B–H bond and it becomes attracted by the large negative electrostatic potential at the bond critical point of one

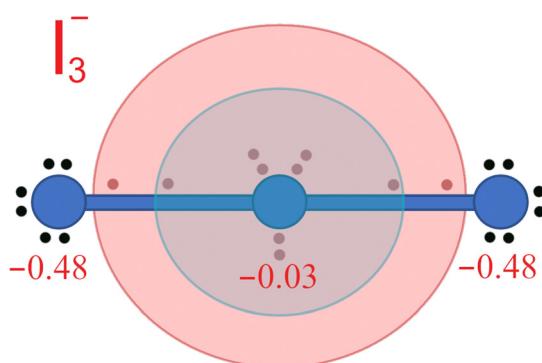


Fig. 5 Schematic distribution of the 22 electrons of the I_3^- polyanion depicted in Fig. 4 when the Bader charges (in red) are considered. Since external atoms have a more negative charge than the central atom, the two electrons present at each collinear I–I bond can be considered to be shifted towards the external atoms. Therefore, electrons around the central atom can be distributed in three spheres. The internal core sphere (dark blue), the intermediate valence sphere (light blue), and the external van der Waals sphere (light red) are shown. ERMB in I_3^- can be interpreted as if 8 electrons were allocated at the valence sphere and 2 electrons were allocated at the van der Waals sphere. In this way, it can be considered that the octet rule is not violated in ERMBs since there are no more than 8 electrons in the valence sphere.



of the B–H bonds of the neighbor BH_3 molecule. Consequently, the two BH_3 molecules approach each other in such a way that there are two simultaneous *trans* influences of two new intermolecular B–H bonds into two intramolecular B–H bonds (see Fig. 4b). As a result, there must be a charge transfer of electrons from the two intramolecular bonds towards the two intermolecular bonds (note that neither B nor H atoms have LEPs that could provide the charge needed to form the two new intermolecular bonds). This charge transfer or electron donation from the intramolecular bonds towards the intermolecular ones leads to the transformation of two covalent B–H bonds (plus their associated weak secondary intermolecular $\text{B}\cdots\text{H}$ bonds) into two B–H–B 3c–2e bonds, as suggested by Walsh.⁶⁴ The charge transfer ends when both the intramolecular and intermolecular bonds become equal and the two B–H–B 3c–2e bonds are formed in the center of B_2H_6 (with all central B–H distances having the same length).

The diborane molecule has four, short, external, ionic-covalent B–H bonds ($d = 2.238 \text{ \AA}$) and four, long, central B–H bonds ($d = 2.478 \text{ \AA}$) which form part of the two B–H–B 3c–2e EDMBs. The larger B–H bond length within the EDMBs than the ionic-covalent B–H bond is explained by the *trans* influence process and accounts for the 0.5 bond order of the B–H bond in the EMB in contrast with the *ca.* 1.0 bond order of the ionic-covalent B–H bond in BH_3 (see Tables S2 and S3, ESI†). The two simultaneous *trans* influences lead to the formation of two supported 3c–2e bonds.¹⁰ These EDMBs are favored because of the filling of the empty sp^3 orbital of the two B atoms, so the B atoms in B_2H_6 reach a more stable tetrahedral geometry than the planar trigonal geometry of the BH_3 molecules (see the molecular orbitals in Fig. S4 of the ESI†).⁶³ Moreover, the octet rule is accomplished for the B atom in B_2H_6 , unlike in the isolated BH_3 molecules, if we take into account for each B atom not only the four electrons of the two ionic-covalent B–H bonds but also the two electrons of each B–H–B 3c–2e bond, as recently proposed.⁶² Note that the 3c–2e bond can be considered within the valence bond theory as a bond in which two electrons resonate between two covalent bonds so that there is one electron in each bond when an average in time is considered.⁷

The fundamental point here is that the mechanism of EMB formation in B_2H_6 is the same as that of EMB formation in the octahedrally-coordinated phases of group-15 and -16 elements, as shown in ref. 1 and 2. In both cases, there is a charge transfer from the primary covalent bond to the secondary non-covalent bond that softens and enlarges the primary covalent bond as shown in Fig. 3d–f. The loss of charge of the primary covalent bond is the reason for the small value of ES in EDMBs once they are formed. The ES value of an EMB is expected to be smaller than the ES value of an ionic-covalent bond of similar polarity (similar ET value) as shown in the green region of EDMBs in Fig. 1. In particular, we have obtained that the ES value of each B–H bond within the 3c–2e EMB to be 0.68 [0.60]. Moreover, the Bader [Löwdin] charge of the different atoms is $\text{B}^{1.80+[1.65+]}$, central $\text{H}^{0.63-[0.59-]}$, and external $\text{H}^{0.59-[0.47-]}$. Since the Bader [Löwdin] charge of the central H atoms is -0.63 $[-0.59]$, this

quantity in absolute value is taken as the ET value for every B–H bond within the EMB. These ES and ET values allow us to locate the central B–H bond as an EMB in the green region of Fig. 1 and Fig. S1 (ESI†), as mentioned in the introduction. In this molecule, the Bader [Löwdin] charge of B and H atoms are $+1.94$ $[+1.72]$ and -0.61 $[-0.57]$. The ET and ES values of the terminal B–H bonds are taken as 0.61 $[0.57]$ and 1.25 $[0.95]$, which allows us to classify these bonds as ionic-covalent ones according to Fig. 1 and Fig. S1 (ESI†), as expected. Note that the short ionic-covalent B–H bond has a value of *ca.* 1.18 \AA which is similar to that experimentally obtained.⁶⁵ Curiously, both the ES and ET values of the EMB in B_2H_6 are similar to those obtained in lead halide perovskites, such as CsPbI_3 .^{25,66} Therefore, our results for B_2H_6 suggest that all EDMBs are located in the green region of Fig. 1 and Fig. S1 (ESI†).

Again, the concepts of the core, valence, and van der Waals spheres can help us now to see that the multicenter interaction (stage 2) starts when $d_2/d_1 = 2$ not only in the formation of ERMBs, such as in the I_3^- anion and the $[\text{FHF}]^-$ molecule, as discussed in the ESI† but also in the formation of EDMBs. We do not have information regarding the EMB formation in B_2H_6 , but we have information regarding hydrogen (H_2) under compression.^{67–70} A discussion on the formation of pressure-induced EDMBs in this system can be found in Section S2.3 of the ESI†.

Since stage 2, when multicenter interaction starts, is crucial to understand the formation of both types of multicenter bonds from mere electrostatic secondary interactions in stage 1, we can make the following statement regarding the formation chemical bonds in general: “it is important to stress that electrostatics (or its extension to electromagnetism), which can be used to describe supramolecular interactions of van der Waals type when $d_2/d_1 > 2$, is not enough to describe the stronger chemical bonds because it only provides a first approximation to the problem. In stronger chemical bonds, such as covalent bonds and multicenter bonds, electrons are so close in real space (closeness measured in terms of the atomic Bohr radius and now in terms of the valence sphere^{47,48}) that quantum mechanics plays a role, as we have shown in stage 2 of the formation of multicenter bonds. Therefore, chemical bonds should be explained with the Schrödinger equation and even this is not always enough. It is well known that properties of materials with heavy elements (with main quantum numbers $n = 5$ and 6, in which valence electrons travel at a fraction of the speed of light) need relativity to be properly explained. Consequently, we conclude that, in the last term, strong chemical bonds should be explained with the Dirac equation that implements the spin-orbit interaction, which also takes into account magnetic interactions. In other words, chemical bonds must be explained within the realm of quantum electrodynamics; *i.e.*, the theory that comprises electromagnetism, quantum mechanics, and relativity.⁷¹ This is the reason why more and more sophisticated computer simulations performing these complex theoretical calculations, mainly developed in the last half a century, have provided us with an ever-increasing power to understand chemical bonds and materials properties in more and more detail.”



It must be noted that the mechanisms of ERMB and E MDB formation in I_3^- and B_2H_6 discussed in this section show only part of the full formation mechanism proposed in Fig. 2 since these examples focus on stage 2 (*trans* influence). This is the stage in which the two mechanisms of multicenter bond formation differ. Now we will provide other examples that illustrate a more complete picture of the mechanism of ERMB and E MDB formation from stage 1 upon increasing the electron density. In most of the examples below, the electron density increases by increasing thermodynamic pressure.

Regarding the ERMB formation, an example of pressure-induced ERMB formation is the pressure-induced hydrogen bond symmetrization due to the formation of infinite zigzag $X-H-X$ chains in hydrogen halides HX ($X = F, Cl, Br, I$).⁷²⁻⁷⁷ HX are simple diatomic molecules forming molecular solids in which a hydrogen bond is found in the condensed state. It has been demonstrated that, at certain conditions of pressure and temperature, a $Cmc2_1$ phase (with infinite zigzag chains showing asymmetric $X-H\cdots X$ bonds) shows a pressure-induced second-order phase transition to a $Cmcm$ phase (with infinite zigzag chains showing symmetric $X-H-X$ bonds) in several HX compounds, as schematized in Fig. 4c for HF. As pressure increases, there is a decrease in the separation of the HX molecules on the left side of Fig. 4c, so both intramolecular and intermolecular bond lengths tend to equalize forming the zigzag chain on the right side of Fig. 4c.⁷²⁻⁷⁷ This is a pressure-induced polymerization process in which the $X-H-X$ bonds are ERMBs as illustrated on the right side of Fig. 4c.

Following Zhang *et al.*,⁷⁷ we have performed simulations of HF that have confirmed the multicenter nature of the $X-H-X$ bonds in the zigzag chain of the $Cmcm$ phase. Simulations for HF show a larger bond length (about 0.2 Å longer) for the $H-F$ bond in the 3c-4e ERMBs at the phase transition pressure (19 GPa) in the $Cmcm$ phase than the covalent $H-F$ bonds present in the molecular $Cmc2_1$ phase of HF at RP (see Table S2, ESI†). This result is similar to that found in previous simulations.⁷²⁻⁷⁷ The equalization of the short $H-F$ and long $H-F$ bonds as pressure increases and the larger $H-F$ bond length at the $Cmcm$ phase than the short covalent $H-F$ bond at $Cmc2_1$ phase evidence the *trans* influence occurring during the compression of the $Cmc2_1$ phase of HF and the multicenter character of the $H-F$ bond in the $Cmcm$ phase.⁷²⁻⁷⁷ The ES and ET values of the covalent $H-F$ bond in the $Cmc2_1$ phase and of the $H-F$ ERMB in the $Cmcm$ phase (see Tables S2 and S3, ESI†) allow us to plot them in Fig. 1 and Fig. S1 (ESI†). Curiously, the ES and ET values of the $H-F$ bonds in the $Cmcm$ phase (see Tables S2, S3 and Fig. S1, ESI† and Fig. 1) suggest that these bonds are closer to E MDBs than to the ERMBs of the $F-H-F$ bonds in the HF_2^- molecule (see Fig. 1). The reason for the decrease of the ES value of the $F-H$ ERMB in the zigzag chain with respect to the isolated HF_2^- molecule is that molecules with ERMBs longer than 3 centers become more and more electron-deficient as the number of bonds increase.^{27,39} This will be discussed further in Section 6. The transformation from a covalent $H-F$ bond plus a secondary hydrogen $H\cdots F$ bond in the $Cmc2_1$ phase to an $F-H-F$ ERMB in the $Cmcm$ phase can be

seen when the 3D isosurfaces of the ELF are plotted (see Fig. S5 of the ESI†). The toroidal isosurface, corresponding to the three LEPs of F atoms, transforms into a cushion-like isosurface corresponding to two LEPs of F atoms since the electrons of one LEP of F atoms now take part in the ERMB in the $Cmcm$ phase. In summary, the symmetric zigzag chain in HX compounds can be considered as a zigzag concatenation of linear 3c-4e ERMBs that already tend to become E MDBs; *i.e.*, they are not pure ERMBs as those of the isolated $X-H-X$ trimers.

Furthermore, the vibrational modes measured and calculated confirm our claims for the multicenter bond formation in the $Cmcm$ phase of HX compounds.⁷²⁻⁷⁷ While low-frequency phonons have a positive pressure dependence, high-frequency phonons show an anomalous softening in the $Cmc2_1$ phase. Instead, all the phonons exhibit a normal positive pressure coefficient when the $Cmcm$ phase is formed.⁷²⁻⁷⁷ Therefore, the example of HF clearly exemplifies the general features observed in Fig. 2 regarding the formation of ERMBs from the structural and vibrational points of view. This example of pressure-induced ERMB formation complements the examples of pressure-induced E MDB formation previously reported in ref. 1 and 2 that will be summarized later.

At this point, we want to comment that the pressure-induced formation of infinite zigzag chains in HX compounds is similar to the imaginary example of pressure-induced ERMB formation due to the polymerization of molecular iodine, as schematized in Fig. 4d. This example has been recently simulated and discussed due to its relationship with the pressure-induced phase transitions in elemental iodine.⁷⁸ In solid I_2 , both intramolecular and intermolecular bond lengths tend to equalize as pressure increases, thus forming the infinite zigzag chain (see Fig. 4d). This can be considered as a zigzag concatenation of linear 3c-4e ERMBs, as those of the linear I_3^- trimer. As for the case of the symmetric zigzag chain in HX compounds, the symmetric zigzag iodine chain in Fig. 4d consists of linked 3c-4e ERMBs that are about 0.2 Å longer than the covalent bonds present in the diiodine molecule⁷⁸ due to the *trans* influence. This comment agrees with our previous suggestion that there is a different bond type in the infinite linear and zigzag iodine chains (Fig. 4d and f).²⁷ As already found for HX compounds in the $Cmcm$ phase, our ES and ET calculations for the symmetric phase of the infinite zigzag iodine chain suggest that the ERMBs in the chain are not pure unlike in the linear I_3^- trimer as can be seen from its place in Fig. 1 and Fig. S1 (ESI†).

Additionally, we can provide one example of ERMB formation due to a reduction of a system (injection of electrons in the system), which also leads to an increase in the electron density of the system. In a recent work on linear 1D-iodine structures encapsulated in single-wall carbon nanotubes, it was shown that the encapsulated iodine atoms tend to form linear I_5^- units and that electron injection in the nanotubes promotes the transformation of the linear I_5^- units into I_3^- units.⁷⁹ We propose that this transformation can be understood as the reaction $3I_5^- + 2e^- \rightarrow 5I_3^-$. It has been already stated regarding Fig. 4a and 5 that I_3^- units are characterized by an ERMB. In addition, it has recently been discussed that linear



I_5^- units do not show ERMBs and are slightly stable linear polyiodide units since isolated I_5^- units tend to be V- or L-shaped.^{24,27} Therefore, the transformation of the linear I_5^- units into I_3^- units upon reduction constitutes a clear example of reduction-induced ERMB formation.

On the other hand, examples of pressure-induced EDMB formation from intermolecular, secondary bonds occur in pnictogens and chalcogens,^{1,2} in the formation of polymeric CO_2 phases,⁸⁰ and in the atomic/polymeric phases of elemental nitrogen⁴² and hydrogen.⁶⁷⁻⁷⁰ The pressure-induced EDMB formation in these systems has recently been discussed.^{1,2} For instance, it has been shown that stage 1 in elemental As occurs between RP and 16 GPa. Above this pressure, the *trans* influence of stage 2 leads to an increase in the primary bond length until both primary and secondary bond distances equalize above 25 GPa (stage 3).^{1,2} The behavior of bond distances, vibrational modes, and the ES value in elemental As as pressure increases (schematized in Fig. 2 and Fig. S3, ESI[†]) can be nicely seen in ref. 1 and 2. However, the most extreme case of pressure-induced EDMB formation likely corresponds to H_2 ,⁶⁷⁻⁷⁰ in which a normal decrease of both intramolecular H–H and intermolecular H···H bond lengths occurs below 100 GPa (stage 1). This is followed by an anomalous increase (decrease) of the intramolecular (intermolecular) bond distance in H_2 as pressure increases above 100 GPa due to the *trans* influence (stage 2). This anomalous increase of the primary bond distance is accompanied by the charge transfer from the intramolecular covalent bonds to the intermolecular non-covalent bonds that ends once both bond distances equalize, and each bond in the atomic/polymeric phase has a single electron per atom pair, *i.e.*, once all H–H bonds become 2c–1e bonds (stage 3), as has been demonstrated in ref. 70. Noteworthy, the charge transfer in elemental hydrogen from the intramolecular covalent bond towards the intermolecular non-covalent bond is out of any doubt since the two electrons of the H_2 molecule at RP are between the two H atoms of the H_2 molecule and there is only one electron per H atom (no LEPs involved). Therefore, it seems that the atomic/polymeric phase (stage 3) of elemental hydrogen, which is supposed to exist above 500 GPa and discussed in ref. 1 and 2, is likely characterized by interacting or multicenter 2c–1e EDMBs that can be considered as extended 3c–2e bonds. Note that the mechanism of EDMB formation shown in Fig. 3d–f is also valid for electron-deficient elements, such as H and B, if the LEP is removed from that figure, as already commented.

The aforementioned stages of the pressure-induced EDMB formation in As and H_2 cannot only be related to the equalization of primary and secondary bond distances, as shown in Fig. 2a; they can also be traced by the behavior of the optical phonon frequencies, as shown in Fig. 2b. In elemental As, there is an anomalous decrease of the stretching phonon frequency from 0 to 25 GPa and an increase above this pressure. These features have been attributed to the weakening (strengthening) of the primary (secondary) bonds and the equalization of both bonds above 25 GPa.^{1,2} In H_2 , the first increase and posterior decrease of the phonon frequencies in the region near 100 GPa

has been attributed to the weakening (strengthening) of the intramolecular (intermolecular) interaction,⁷⁸ which clearly can be ascribed to the charge transfer (*trans* influence) from the intramolecular bonds to the intermolecular bonds, as already commented in ref. 70.

Noteworthy, the case of EDMB formation in elemental hydrogen at HP is similar to the case of the infinite linear chain of hydrogen atoms (Fig. 4e), which in turn is similar to the case of the infinite linear chain of iodine atoms (Fig. 4f). The only possible electron distribution shown in Fig. 4e and f, when the VSEPR theory is considered,⁸¹ clearly shows that these two examples of hydrogen and iodine polymerization show EDMBs in both electron-deficient and electron-rich elements, respectively, *i.e.*, in elements without LEPs and with LEPs, respectively. In both cases, the intramolecular bond of the hydrogen (H_2) and iodine (I_2) molecules at low pressure loses its covalent character as a 2c–2e bond upon polymerization when the molecules approach each other at increasing pressure. The *trans* influence of the secondary intermolecular bond into the primary intramolecular bond (as both tend to equalize distances with increasing pressure) forces a charge transfer from the primary bond towards the secondary bond that results in a final electron-deficient 2c–1e bond, irrespective of the electron-deficient or electron-rich character of the bonded elements, thus resulting in the formation of EDMBs in the infinite linear atomic chains.

To conclude this section, we want to stress that the different mechanisms of ERMB and EDMB formation presented in this section and the ES and ET values found for different compounds with ERMBs and EDMBs allow us to reaffirm our conclusions in ref. 1 and 2; *i.e.*, that bonds in crystalline PCMs and the octahedrally-coordinated A_h and A_i phases of group-15 and -16 elements are indeed EDMBs. These EDMBs are similar to those of boranes and also similar to those occurring in elemental nitrogen and hydrogen at HP once intermolecular and intramolecular bonds equalize. Therefore, we can conclude that EDMBs can be observed in materials with both electron-deficient and electron-rich elements. This is a change of paradigm since EDMBs were supposed not to exist in electron-rich elements and will be further discussed in Section 6. The EDMB formation is characterized by the coexistence of localized and delocalized electrons and is an intermediate step between ionic-covalent bonding, with fully localized electrons, and metallic bonding, with fully delocalized electrons. This explains the intermediate position of the EDMB between the covalent and metallic bonds in the ES *vs.* ET maps in Fig. 1 and Fig. S1 (ESI[†]).

It should be stressed that our view of EDMBs and their position in the ES *vs.* ET maps is consistent with the well-known progressive delocalization of electrons as pressure increases until the metallic bond is reached at enough HP.^{39,42} It is significant that the occurrence of EDMBs under compression, as in the A_h and A_i phases of group-15 and -16 elements, in atomic/polymeric nitrogen and hydrogen, and, in general, in all materials at HP, agrees with the already expressed general view that all elements, irrespective of their valence electrons, should show multicenter bonds (it should be



interpreted in many cases as EDMBs) at HP.^{39,82} Note that pressure tends to increase atomic coordination in all materials, thus all atoms, sooner or later above a given atomic coordination, will fall short in electrons to share two electrons with neighbor atoms.⁴² Therefore, all the new bonds in those pressure-induced hypercoordinated atoms will have to share necessarily less than two electrons per atomic pair until full electron delocalization, typical of metallic bonding, is finally attained at very HP, even in hydrogen.⁸³

All in all, in this section we have shown, using the examples of the I_3^- and B_2H_6 molecules, the examples of EDMBs in pnictogens and chalcogens of ref. 1 and 2, and the examples of ERMBs in the symmetric zigzag chains of HF and I_2 , that the mechanisms of ERMB and EDMB formation are different and that the ES and ET values allow us to distinguish between EDMBs and ERMBs thanks to the ES vs. ET maps obtained either with density-based or orbital-based methods. Our different ES values for ERMBs and EDMBs agree with recent works that have shown the different ES values for molecules with ERMBs (XeF_2 , ClF_3 , and SF_4) and for solids with EDMBs (octahedrally coordinated crystalline phases of pnictogens, chalcogens, and PCMs).^{2,25}

4. ERMBs and EDMBs in materials with electron-rich elements

To prove the change in paradigm related to multicenter bonds in electron-rich elements, we provide in this section examples of materials made of electron-rich elements in which there are linear ERMBs and EDMBs in 1D, 2D, and 3D. In particular, we are going to comment on examples that were previously assumed to be ERMBs²⁰ and they are indeed EDMBs. As done in the previous section, to evidence the presence of the two kinds of multicenter bonds and to distinguish between them, we will evaluate the bond distances, the Bader [Löwdin] atomic charges, as well as the ES and ET values between two atoms in the different bonds present in these systems (see data in Tables S2 and S3, ESI[†]).

As aforementioned, EDMBs and ERMBs have been considered until now to be present only in molecules and solids with electron-deficient and electron-rich elements, respectively.⁸⁴ However, it is curious that hypervalent bonds (or ERMBs) in 1D, 2D, and 3D present in solids with electron-rich elements, such as Li_2Sb , $BaZnSb_2$, and sc-Sb, respectively, have been traditionally considered to share less than two electrons between two atoms.²⁰ In other words, the assumed ERMBs in these solids have been considered to be 2c-1e bonds according to the Pimentel model,^{18,19} like the EDMBs of B_2H_6 and sc-Po.

Despite assuming that the bonds in sc-Sb, isostructural to sc-Po,⁸⁴ are 2c-1e bonds, these bonds have not been considered to be EDMBs but ERMBs,²⁰ being the only distinction between ERMBs and EDMBs the assumption that ERMBs and EDMBs are found in molecules and solids with electron-rich and electron-deficient elements, respectively.⁸⁴ This argument to distinguish between EDMBs and ERMBs is at odds with the

number of electrons shared (ES) and the renormalized number of electrons transferred (ET) between two atoms for these two types of multicenter bonds according to Fig. 1 and Fig. S1 (ESI[†]), as we have reasoned in the previous section.

In a recent work,²⁷ we have mentioned that part of the misunderstanding in distinguishing between ERMBs and EDMBs comes from the simple molecular diagram used to understand 3c-2e EDMBs and 3c-4e ERMBs; *i.e.*, the Pimentel three-level model.^{18,19} According to this diagram (see Fig. 1 in ref. 27), it is thought that the two electrons occupying the non-bonding orbital in 3c-4e bonds belong exclusively to the external atoms and are not even partially shared with the central atom. However, this reasoning results in a contradiction of the hypervalent model because if those electrons are not shared with the central atom, then there is no violation of the octet rule for the central atom of the trimer and there should be no hypervalency. On the contrary, if it is assumed that there is a violation of the octet rule that leads to hypervalency then it must be assumed that these electrons are fully shared with the central atom, so the molecular diagram is not valid and the ERMBs cannot be considered as 2c-1e bonds. The contradiction can be solved if we consider that ERMBs are not 2c-1e bonds and that the classical molecular diagram of Fig. 1 in ref. 27 is not valid to understand 3c-4e bonds. This is because the electrons of the non-bonding orbital are partially shared (not fully shared) with the central atom in a trimer as the rather large ES values of ERMBs suggest. Note that the partial sharing of these electrons explains the smaller ES value in ERMBs than in covalent bonds, the larger ES value in ERMBs than in EDMBs (since ERMBs are not 2c-1e bonds unlike EDMBs), and the lack of a severe violation of the octet rule for the central atom in the ERMB (since the extra electrons in 3c-4e bonds with respect to 3c-2e bonds are only partially shared with the central atom as if they would be in a secondary electronic valence sphere, the van der Waals sphere,^{47,48} as explained for the I_3^- anion in Fig. 5).

4.1. Solids with ERMBs

In this subsection, we provide examples of different materials with electron-rich elements and well-known ERMBs. Examples of 1D ERMBs are the I_3^- polyanion (Fig. 4a and 5) alone or inside solid CsI_3 , the HF_2^- polyanion, and the XeF_2 molecule. Examples of 2D ERMBs are the XeF_4 molecule and the planar Te_5^{2-} molecules in solid Cs_2Te_5 . Examples of 3D ERMBs are the SF_6 and XeF_6 molecules and the square TeI_6^{2-} molecules in solid Cs_2TeI_6 . Since 3c-4e ERMBs in molecules with electron-rich elements have been thoroughly discussed in the literature, we will focus here only on the three mentioned solids. The crystalline structures of these solids (CsI_3 , Cs_2Te_5 , and Cs_2TeI_6) are shown in Fig. 6 and the relevant data are provided in Tables S1-S3 (ESI[†]).

As can be seen in Fig. 6 and Tables S2 and S3 (ESI[†]), the I_3^- , Te_5^{2-} , and TeI_6^{2-} molecules in CsI_3 , Cs_2Te_5 , and Cs_2TeI_6 exhibit ERMBs with larger Bader [Löwdin] charges in the external atoms than in the central atom of the 3c-4e bonds and larger values of ES in the corresponding two-center bonds within the molecule with ERMBs than those found in the



corresponding two-center bonds in solids with EDMBs for similar ET values.^{2,25} Consequently, ERMBs in these solids are located at the orange color region of Fig. 1 and Fig. S1 (ESI†) slightly above the red region of covalent bonds and much above the green region of EDMBs. These results for solids agree with the results we have previously shown for I_3^- , HF_2^- , and XeF_2 molecules with ERMBs and with previous simulations on similar molecules (XeF_2 , ClF_3 , and SF_4).²⁵

Let us comment in more detail on the results of ERMBs in the aforementioned solids. In CsI_3 (Fig. 6a), all the I–I bonds are more than 0.1 Å larger than the covalent I–I bond in the *Cmca* phase of elemental iodine at RP.⁸⁵ The ES and ET values of all the I–I bonds in CsI_3 are similar to those in the I_3^- polyanion (see Tables S2 and S3, ESI†), so they can be classified as ERMBs and are located close to the I_3^- polyanion in the ES vs. ET map (see Fig. 1 and Fig. S1, ESI†). For the I_3^- polyanion (alone or in CsI_3), the ET value is taken as the difference in the Bader [Löwdin] charge between the terminal and the central I atoms of the I_3^- polyanion (note that the charge of the central atom must be divided between the two terminal atoms). Similarly, all the I–Te–I 3c–4e bonds in Cs_2TeI_6 (Fig. 6c) show I–Te bond distances of 2.925 Å, which are much larger than ionic-covalent I–Te bonds with an average bond distance of *ca.* 2.74 Å in monoclinic TeI_4 .⁸⁶ Note that the I–Te bond in Cs_2TeI_6 exhibits ES and ET values of 1.42 [0.94] and 0.52 [0.34], respectively, so they can be classified as ERMBs and are plotted correspondingly in Fig. 1 and Fig. S1 (ESI†). These values are in contrast with the ES and ET values of *ca.* 1.90 [1.4] and 0.15 [0.13] that correspond to ionic-covalent I–Te bonds in monoclinic TeI_4 with calculated distances *ca.* 2.77 Å (see Table S2, ESI†). We want to stress that obtaining the ET value in ternary compounds could be very challenging because it could be difficult to know between which atoms the electronic charge has been transferred. For Cs_2TeI_6 , we have taken as ET the absolute value of the difference in Bader [Löwdin] charge between the external I atom and the central Te atoms of the TeI_6^{2-} polyanion (taking into account that the charge of the central atom must be divided between the six terminal atoms).

A different case occurs for Cs_2Te_5 (Fig. 6b). This solid contains two distinct types of Te–Te bonds. The Te(16h)–Te(16h) bonds

mainly along the *b*-axis are short ($d = 2.82$ Å) and have ES = 2.2 [1.65] and ET = 0 [0], so they can be classified as pure covalent bonds. On the other hand, the long Te(4c)–Te(16h) bonds ($d = 3.06$ Å) within the planar Te_5^{2-} array extended in the *ac* plane have ES = 1.43 [0.94] and ET = 0.39 [0.33], so they can be classified as ERMBs. It must be noted that the molecular compounds with ERMBs usually have large bandgaps, like in CsI_3 ⁸⁷ and Cs_2TeI_6 ,⁸⁸ however, low bandgaps and relatively high electrical conductivity were reported in Cs_2Te_5 along the layer plane (likely due to the zigzag connectivity of ERMBs and covalent bonds along the layer plane).⁸⁹

In the above solids, it can be considered that the ERMB formation in the I_3^- , Te_5^{2-} , and TeI_6^{2-} molecules is promoted by the increase in their electron densities, which is caused by the reduction of these systems due to the donation of electrons from the Cs atoms. In other words, Cs atoms give their electrons to the I_3 , Te_5 , and TeI_6 polyanions to form nominal I_3^- , Te_5^{2-} , and TeI_6^{2-} molecular units. These molecular units exhibit 3c–4e ERMBs in 1D, 2D, and 3D (in which the central atom is hypercoordinated) and are linked to Cs atoms by ionic bonds. This view is consistent with the Bader [Löwdin] atomic charge expressed as a superindex in the three compounds; *e.g.* CsI_3 at RP has $Cs^{0.73+[0.65^+]}$. Note that in the I_3^- units of CsI_3 at RP there are two slightly different I–I bonds (one short and one long) that tend to form two equal ERMBs at HP,⁹⁰ *i.e.*, equal bonds as shown by the linear array of the I_3^- polyanion in Fig. 4a. A similar case occurs for Cs_2Te_5 (see Fig. 6b) with two $Cs^{0.70+[0.67^+]}$ atoms giving their charges to the five Te atoms of the Te_5 molecule to form the Te_5^{2-} polyanion. In this case, there are two independent Te atoms in the Te_5^{2-} polyanion that show two different negative charges with the charge of the external Te atoms being larger than that of the central Te atom, as expected for 3c–4e bonds. In particular, Te(4c) and Te(16h) atoms are $Te^{0.03+[0.01^+]}$ and $Te^{0.36-[0.34^-]}$, respectively. The different atomic charges of the central (positive) and external (negative) Te atoms in the planar Te_5 array are consistent with the ERMB nature of these bonds in 2D. Finally, the Bader [Löwdin] atomic charges in Cs_2TeI_6 are: $Cs^{0.76+[0.63^+]}$, $Te^{0.82+[0.68^+]}$, and $I^{0.39-[0.32^-]}$. This means that the TeI_6^{2-} units (see Fig. 6c) have a central Te atom that also gives almost one

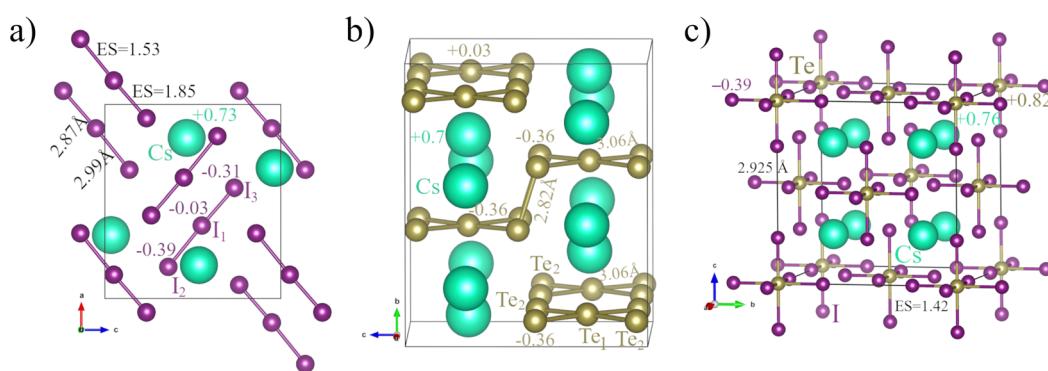


Fig. 6 The crystal structure of solids (a) CsI_3 , (b) Cs_2Te_5 , and (c) Cs_2TeI_6 . These solids show molecular units with 3c–4e bonds (ERMBs) in 1D, 2D, and 3D. In CsI_3 , I_1 , I_2 , and I_3 atoms correspond to three independent 4c sites. In Cs_2Te_5 , Te_1 and Te_2 atoms correspond to Te(4c) and Te(16h) atoms, respectively. Details on bond distances and Bader atomic charges are illustrated here and also summarized in Table S2 (ESI†).



electron to the terminal I atoms to form ERMBs in 3D. Therefore, as expected for linear 3c–4e ERMBs (in this case in three perpendicular directions), the terminal I atoms have a larger electronic charge, so this cubic TeI_6^{2-} unit behaves as a pseudo- SbI_6^{3-} unit that is isoelectronic to XeF_6 . Note that the central atoms of the 3c–4e bonds (*e.g.* Te(4c) in Cs_2Te_5 and Te in Cs_2TeI_6) have positive Bader [Löwdin] charges, so they behave as cations.

In brief, the three mentioned solids in this subsection are Zintl phases that can be understood on the light of the Zintl–Klemm concept by which atoms can exchange electronic charge irrespective of their electronegativity or electron affinity values.^{20,91,92} The charge donated by Cs atoms is accepted by the polyanions in such a way that Cs atoms are linked to the polyanions *via* ionic bonds while there are homonuclear or homoatomic bonds (bonds between the same atomic species, Te–Te and I–I) and heteronuclear or heteroatomic bonds (bonds between different atomic species, Te–I) some of them being ERMBs inside the polyanions.

4.2. Solids with EDMBs

Let us discuss the formation of EDMBs in 1D, 2D, and 3D in solids with electron-rich elements. Here we are going to present the solids with EDMBs in two main categories: solids with homopolar (or homonuclear or homoatomic) EDMBs (Fig. 7) and solids with heteropolar (or heteronuclear or heteroatomic) EDMBs (Fig. 8). In the first category, we will comment on the formation of (i) 1D EDMBs in the infinite linear Sb chains in Li_2Sb , (ii) 2D EDMBs in the square planar array of Sb atoms in solid BaZnSb_2 , and (iii) 3D EDMBs in the octahedrally-coordinated phases of pnictogens and chalcogens, such as the sc phase of Po (α -Po) at RP, which is isostructural to the HP phases sc-As and sc-Sb. In the second category, we will comment on the formation of (i) 1D EDMBs in TeO_2 at RP (also the cases of SbPO_4 , and $\text{Sc}_2\text{Si}_2\text{O}_7$ at RP are discussed in the ESI†), (ii) 2D EDMBs in the *Cmcm* phase of SnSe at HP (also of GeSe at HP), which is isostructural to the low-pressure phase of InBr, InI, and TlI, and (iii) 3D EDMBs in the rocksalt (rs) phase of β -GeTe at HP, isostructural to PbS, PbSe, and PbTe at RP.

4.2.1. Homopolar EDMBs. Homopolar EDMBs are characterized by being EDMBs with $ET = 0$, so they are typically formed by a single element; *i.e.*, they are usually homoatomic bonds. A clear example is the case of bonds between atoms that occupy the same Wyckoff site. The first example of solid with homopolar EDMBs we are going to comment is solid sc-Po (α -Po), isostructural to sc-As and sc-Sb, in which 3D EDMBs are observed (Fig. 7c).^{1,2} In sc-Po, Po–Po EDMBs occur in the three spatial directions with $ES (ET) \approx 1 (0)$ at RP (see Fig. 1). The formation of EDMBs in sc-Po at RP and in sc-As and sc-Sb at HP as well as in the rhombohedral (rh) phase of Po (β -Po) at RP and of rh-Se and rh-Te at HP have been thoroughly discussed in ref. 1 and 2. Note that the rhombohedral β -Po phase is a slightly distorted modification of the cubic α -Po phase which has been found in all chalcogens at different pressures. Noteworthy, our results are contrary to a previous work in which the α -Po phase

and the same phase in the other pnictogens (sc-Sb) were previously assumed to exhibit 3D ERMBs.²⁰

The second example of homopolar EDMB is the solid BaZnSb_2 (Fig. 7b) in which 2D EDMBs can be observed. This is an intermetallic compound that was also previously considered to feature 2D $\text{Sb}_1\text{–Sb}_1$ ERMBs.²⁰ We find, in a similar way as for α -Po (and β -Po),^{1,2} that the infinite square planar array of Sb_1 atoms in BaZnSb_2 features $\text{Sb}_1\text{–Sb}_1$ bonds with $ES = 1.16 [0.8]$ and $ET = 0 [0]$. The important point here is that these ES and ET values are closer to those of EDMBs in α -Po, crystalline PCMs, such as β -GeTe and SnTe, and B_2H_6 than to those of ERMBs (see Fig. 1 and Fig. S1, ESI†).[¶] Note that the $\text{Sb}_1\text{–Sb}_1$ bond distance in BaZnSb_2 at RP ($d = 3.24 \text{ \AA}$) is of the same order as that of the Sb–Sb bonds reported in sc-Sb at RP ($d = 3.16 \text{ \AA}$)⁹³ and much larger than the more covalent Sb–Sb bond distance of Sb in the A7 phase at RP ($d = 2.96 \text{ \AA}$).^{1,2} Therefore, we conclude that $\text{Sb}_1\text{–Sb}_1$ bonds in BaZnSb_2 at RP are EDMBs with a half or partial bond order. It is important to mention that our claim for EDMBs between Sb_1 atoms of the square planar Sb array in BaZnSb_2 is supported by Jeitschko and coworkers, who also suggested the presence of EDMBs with 0.5 bond order in As–As and Sb–Sb bonds in the square planar array of isostructural intermetallic compounds ACuAs_2 and AAgSb_2 (A = rare earth and uranium).⁹⁴ Moreover, our claim for the presence of EDMBs in BaZnSb_2 is also supported by Nesper, who suggested that Bi–Bi bonds at the square planar array of Bi atoms in the Zintl phase of LiBi are 2c–1e bonds with 0.5 bond order.⁹⁵ Our results for pnictogens and chalcogens^{1,2} and our current theory of multicenter bond formation provide a clear explanation for the results of these two works.

The third example we have chosen is solid Li_2Sb , in which EDMBs are found in 1D (Fig. 7a). This compound crystallizes in space group no. 189⁹⁶ and have two independent Sb atoms located at 1b and 2c sites. Again, as for the cases of sc-Sb and BaZnSb_2 , both $\text{Sb}_1\text{–Sb}_1$ and $\text{Sb}_2\text{–Sb}_2$ bonds of Li_2Sb ($d = 3.260 \text{ \AA}$ in both cases) were previously assumed to be 1D ERMBs.²⁰ However, the ES and ET values for the two types of Sb–Sb bonds in Li_2Sb at RP are $1.41 [0.7]$ and $0 [0]$, respectively. These values allow us to locate these Sb–Sb bonds in the green region of EDMBs in Fig. 1 and Fig. S1 (ESI†). Note that Sb–Sb bonds in Li_2Sb are even longer than the EDMBs commented in the previous paragraph for BaZnSb_2 and sc-Sb at RP. Therefore, we conclude tentatively that the Sb–Sb bonds in Li_2Sb are EDMBs and not ERMBs unlike previously assumed.²⁰ Our results are thus consistent with the suggestion of 2c–1e bonds present in the 1D infinite linear Sb chains in U_2TiSb_3 and isostructural antimonides by Jeitschko and coworkers.⁹⁷

To close this subsection, we want to comment that the Sb–Sb EDMBs in the studied solids sc-Sb (3D), BaZnSb_2 (2D), and Li_2Sb (1D) have ES values of around $1.0 [0.9]$, $1.2 [0.8]$, and $1.4 [0.7]$, respectively. In other words, the ES values of the Sb–Sb

[¶] Note that in the *I4/mmm* structure of BaZnSb_2 there are two different Wyckoff sites for Sb atoms (4c and 4e sites). Each independent Sb atom has a different charge. All Sb atoms in the square planar array occupy the 4c sites, so $ET = 0$ for $\text{Sb}(4c)\text{–Sb}(4c)$ bonds.



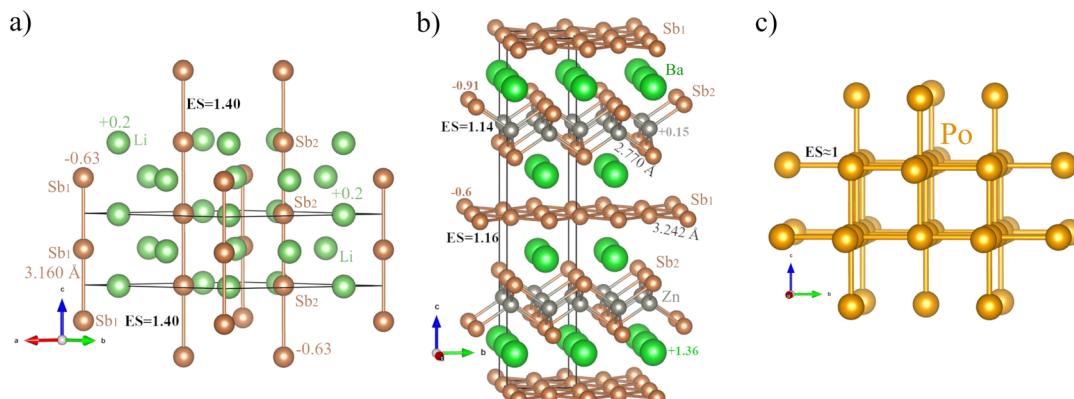


Fig. 7 The crystal structure of solids (a) Li_2Sb , (b) BaZnSb_2 , and (c) sc-Po (α -Po). These solids show hypercoordinated units with EDMBs in 1D, 2D, and 3D. In Li_2Sb , Sb_1 and Sb_2 atoms correspond to $\text{Sb}(1\text{b})$ and $\text{Sb}(2\text{c})$ atoms, respectively. In BaZnSb_2 , Sb_1 and Sb_2 atoms correspond to $\text{Sb}(4\text{c})$ and $\text{Sb}(4\text{e})$ atoms, respectively. In Li_2Sb , Sb atoms form hypercoordinated units with 1D EDMBs and a linear geometry as that of Fig. 10a. In BaZnSb_2 , Sb atoms form hypercoordinated units with 2D EDMBs and a square planar geometry as in Fig. 10b. In sc-Po, Po atoms form hypercoordinated units with 3D EDMBs and a cubic geometry as that of Fig. 10c. Details on bond distances, Bader atomic charges, ET, and ES of various bonds are illustrated here and summarized in Table S2 of the ESI.†

bonds increase as we move from 3D to 1D. At present, we do not yet definitely know the cause of this deviation, but we think that the reason could be related to the Bader [Löwdin] charge of the Sb atoms in those bonds. These charges for Sb atoms are 0 [0], -0.6 [-0.8], and -1.63 [-1.0] for sc-Sb, BaZnSb_2 , and Li_2Sb , respectively. This means that while Sb atoms in sc-Sb behave as neutral Sb atoms (Sb^0), the Sb_1 atoms in the Zintl compound BaZnSb_2 (ideally Sb^-) behave as pseudo-Te atoms, and the Sb atoms in the Zintl compound Li_2Sb (ideally Sb^{2-}) behave as a pseudo-I atoms. In other words, the larger the atomic charge of the Sb atom the larger the ES value. A possible explanation for this behavior is the difficulty of the nucleus of Sb to retain close to the nucleus the extra electronic charge given by the most electropositive atoms, so this extra electronic charge is more shared with the neighbor atoms thus leading to a larger ES value. Another possible explanation is that this trend is related to the electron count: in 3D systems (sc-Sb), each Sb atom participates in three EDMBs (six 2c-1e bonds), requiring three electrons for bonding. In 2D systems (BaZnSb_2), each Sb atom participates in two EDMBs (four 2c-1e bonds), thus requiring

fewer electrons (ideally two) for bonding. Finally, in 1D systems (Li_2Sb), even fewer electrons (ideally one) are required to form one EMB (two 2c-1e bonds). Therefore, the availability of electrons in Sb atoms for bonding through EDMBs is larger as the dimensionality of the EMB decreases – although the extra electrons that do not take part in EDMBs are assumed to form part of LEPs and consequently not to participate in bonding –, so perhaps a small contribution of these extra electrons is reflected in the ES of the EDMBs.

4.2.2. Heteropolar EDMBs. Heteropolar EDMBs are characterized by being EDMBs with $\text{ET} \neq 0$, so they are typically formed by two different elements; *i.e.*, they usually constitute heteroatomic bonds. The first example we are going to comment on is that of β -GeTe, a crystalline PCM in which 3D EDMBs formed by Ge and Te can be observed (Fig. 8c). In β -GeTe at 5 GPa,⁹⁸ Ge and Te atoms occupy the 4a and 4b Wyckoff sites of the rs structure and all Ge-Te bonds ($d = 2.892 \text{ \AA}$) have a calculated ES and ET values of 1.16 [1.11] and 0.18 [0.09]. Here the ET value is calculated by dividing the absolute Bader [Löwdin] charge of Ge or Te atoms by the absolute nominal

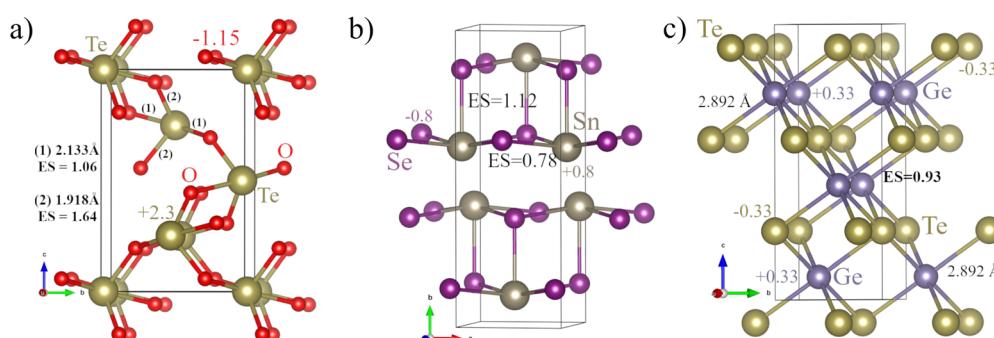


Fig. 8 The crystal structure of solids (a) TeO_2 , (b) Cmcm-type SnSe , and (c) β -GeTe. These solids show heteropolar 1D, 2D, and 3D EDMBs, respectively. TeO_2 shows hypercoordinated units with a see-saw geometry as that of Fig. 10a. Cmcm-type SnSe shows hypercoordinated units with a square pyramidal geometry as in Fig. 10b. β -GeTe shows hypercoordinated units with a cubic geometry as that of Fig. 10c. Details on bond distances, Bader atomic charges, ET, and ES of various bonds are illustrated here and summarized in Table S2 of the ESI.†



charge of Ge and Te atoms (2) since both Ge and Te atoms have the same atomic multiplicity.|| According to the ES and ET values, Ge–Te bonds, which are considered as metavalent bonds by Wuttig and collaborators,^{2,14} can be classified as EDMBs (see Fig. 1 and Fig. S1, ESI†). Note that we get a value of density-based ET slightly smaller than that obtained by Wuttig and coworkers (~ 0.23).^{2,14} It is also interesting to note that the EDMBs in β -GeTe are dative (or coordinative) EDMBs in a similar way as there are covalent and dative covalent bonds. This means that in β -GeTe, unlike in sc-As, in which all As atoms contribute equally to the EMB, Te atoms contribute more to the EMB than Ge atoms because Te atoms have more valence electrons than Ge atoms. Note that GeTe is isoelectronic to As since both are 10-electron materials. In order to establish 3D EDMBs in sc-As, each As atom contributes with three p-type electrons (one along each direction) since the two s-type electrons are non-bonding (forming part of an inactive LEP). In β -GeTe, the two s-type electrons of Ge and Te are also non-bonding (forming part of inactive LEPs), so there are two p-type electrons in Ge and four p-type electrons in Te to form 3D EDMBs. This means that Ge lacks one p-type electron to form an EMB along one direction while Te has one extra p-type electron to form an EMB. As a result, Te gives its extra electron to Ge to form the EMB along one direction, so on average both Ge and Te have three p-type electrons, like As. This allows both atoms to form EDMBs along the three perpendicular directions. In other words, dative or coordinative EDMBs occur in β -GeTe.

The second example we want to show is the layered orthorhombic *Cmcm* (also noted as *Bbmm*) phase of both GeSe and SnSe at HP (see Fig. 8b).^{99–101} Both GeSe and SnSe crystallize in the layered orthorhombic *Pnma* phase at RP. In the *Pnma* structure, both Ge (Sn) and Se atoms are threefold coordinated *via* dative iono-covalent bonds, whereas in the *Cmcm* phase all atoms are fivefold coordinated. The *Cmcm* phase of these compounds at HP is isostructural to that of InBr, InI, and TlI at RP.¹⁰² It has been reported that there are two different covalent bonds in the *Pnma* phase of SnSe and GeSe, and that both bonds evolve differently under compression: the bond almost perpendicular to the layers (axial bond) and the two equivalent bonds almost parallel to the layers (equatorial bonds).^{99–101} Whereas the axial bond evolves as a normal iono-covalent bond since the bond length decreases under compression and its related high-frequency vibrational modes show a positive pressure coefficient, the equatorial bonds show anomalous behavior under compression with an increase of the bond length and a softening of related vibrational modes at HP. Our simulations of SnSe at HP reproduce nicely the behavior of bond lengths experimentally reported (see Fig. S6 in ESI†).⁹⁹ The different behavior of both iono-covalent bonds at the *Pnma* phase can be understood if we consider that the axial bond

remains as a short iono-covalent bond when the *Cmcm* phase is attained at HP whereas the equatorial bonds suffer a *trans* influence that ends with the formation of two long EDMBs in the layer plane. In other words, in the *Cmcm* phase of SnSe and GeSe, the quasi-linear 2D EDMBs can be observed perpendicular to the *b*-axis (see Fig. 8b). The iono-covalent (EDMB) nature of the axial (equatorial) bonds at the *Cmcm* phase of SnSe are confirmed by our ES and ET values (see Table S2 and Fig. S6, ESI†) according to the location in Fig. 1 and Fig. S1 (ESI†). A detailed explanation of the ES and ET values of the *Cmcm* phase of SnSe at 10 GPa and its comparison with the *Cmcm* phase of TlI at RP is provided in Section 2.7 of the ESI.†

The third example we want to comment on is the solid paratellurite (α -TeO₂), in which quasi-linear 1D Te–O EDMBs are formed. In paratellurite at RP, Te and O atoms are at 4a and 8b sites, respectively. Each Te atom is fourfold coordinated to O atoms with two short ($d = 1.919$ Å) and two long ($d = 2.087$ Å) Te–O bonds in a see-saw geometry (note that Te⁴⁺ has a single LEP) derived from the trigonal bipyramidal geometry (Fig. 8a).¹⁰³ The two short (long) bonds have ES = 1.64 [1.5] (1.06 [0.9]) and both have ET = 0.29 [0.23]. Note that to calculate the ET value for the Te–O bonds we have to take the Bader [Löwdin] charge of the Te atom and divide it by the nominal valence of Te (+4). In addition, this value must be divided by the atomic multiplicity ratio between Te and O atoms (2), which is the average number of O atoms to which the electrons of Te atoms are transferred. Note that, equivalently, the ET value can be calculated by taking the absolute Bader [Löwdin] charge of the O atom and dividing it by the nominal valence of O (−2) and also by the multiplicity ratio between Te and O (2). The first procedure can be used more frequently since there are usually more anions than cations in many materials; however, the last procedure can be used when there is more than one single cation providing charge to a single anion. With the ES and ET values we can see that the short Te–O bonds can be considered iono-covalent bonds, while the long Te–O bonds are EDMBs (see Fig. 1 and Fig. S1, ESI†). Note that the smaller ES value of the long bonds than of the short bonds (both with same ET value) necessarily implies that, if the two short bonds are iono-covalent, the two long bonds must be EDMBs since ERMBs are expected to have larger values of ES than the iono-covalent bonds for a given ET as seen in Fig. 1 and Fig. S1 (ESI†). Support to our conclusions about the bonds in paratellurite comes from the compounds BaTeO₃¹⁰⁴ and CoTeO₃.¹⁰⁵ In both compounds, Te atoms are linked to three O atoms with only one type (iono-covalent) of short Te–O bonds and with lengths of the order of 1.86–1.88 Å and 1.90–1.92 Å, respectively. These bonds have similar bond lengths to those of short iono-covalent bonds in α -TeO₂, thus giving support to the EMB nature of the long Te–O bonds in paratellurite.

An additional example of quasi-linear 1D EMB is that formed by Sb and O bonds in SbPO₄ (see explanation in Section S2.8 of the ESI† regarding Fig. S7a).¹⁰⁶ At this point, it is interesting to comment that the calculation of ET values in heteronuclear bonds is easy for binary compounds with a single Wyckoff site for each element and the same multiplicity, such as in β -GeTe or SnSe. In this case, the charge lost by one atom is

|| In this work atomic multiplicity refers to crystallographic multiplicity. In crystallography, the atom multiplicity is the number of the Wyckoff site of an atom. This is different from the chemical multiplicity which is the number of possible orientations of the total spin relative to the total orbital angular momentum.



the charge gained by the other and the normalized number of transferred electrons can be readily calculated by the Bader [Löwdin] charge divided by the nominal atomic valence. However, it is more complex to calculate ET when the charge is gained or lost by more than one element, as already commented. We have shown the cases of TeO_2 where the charge of each Te atom is transferred to two O atoms and the case of a ternary compound (SbPO_4) in which the charge of two atoms Sb and P is transferred to two O atoms on average. In these complex cases, the normalized ET value must be calculated by additionally dividing the previous value (Bader [Löwdin] charge divided by the nominal valence) by the multiplicity ratio between the donor and acceptor atoms. In this way, the different bond types could be reasonably located in Fig. 1 and Fig. S1 (ESI†).

Up to this point, we have suggested that ERMBs and EDMBs can be found in molecules and solids with electron-rich elements, like As, Sb, Te, and I. In fact, EDMBs have also been recently suggested to occur in iodates at different pressures.¹⁰⁷ However, we must mention that we are not the first authors to propose that EDMBs could be observed in electron-rich elements. Vegas and collaborators already suggested that 3c–2e bonds should be present in the linear Si–C–Si bond of the carbocation $[\text{Si}_2(\text{CH}_3)_7]^+$, in the linear Si–O–Si bonds of hexamethyldisiloxane, $(\text{H}_3\text{C})_3\text{Si}–\text{O}–\text{Si}–(\text{CH}_3)_3$, as well as in the linear Si–O–Si bonds within the $[\text{O}_3\text{Si}–\text{O}–\text{SiO}_3]^{6-}$ polyanion in solid $\text{Sc}_2\text{Si}_2\text{O}_7$ silicate.^{92,108} In addition, it must be mentioned that electron-deficient bonds have been proposed to exist in several large C–C bonds, such as those present in carboranes and hydrocarbons,^{65,109} in large O–O bonds in H_2O_2 , and in weak F–F bonds in F_2 . Some of these bonds were classified as protocovalent, charge-shift bonds, and electron-deficient covalent bonds.^{110–113} However, there is a large confusion regarding these systems and the notation of their bondings since charge-shift bonds have been clearly shown to be related to ERMBs (which are not electron-deficient bonds), like XeF_2 .¹¹⁴ Note also that relatively strong secondary bonds between different molecules, in which LEPs are involved in a dative or donor way as we propose for ERMBs, have been also classified as dative or coordinate covalent bonds.^{115,116} To illustrate one of the above examples, we have performed calculations for monoclinic $\text{Sc}_2\text{Si}_2\text{O}_7$ at RP (see explanation in Section S2.8 of the ESI† regarding Fig. S7b) and found two types of Si–O bonds within the $[\text{O}_3\text{Si}–\text{O}–\text{SiO}_3]^{6-}$ polyanion. The terminal (central) heteropolar Si–O bonds of the $[\text{O}_3\text{Si}–\text{O}–\text{SiO}_3]^{6-}$ polyanion can be classified as iono-covalent bonds (EDMBs), thus confirming the suggestion of Vegas and collaborators.^{92,108}

To finish this section, it must be stressed that the presence of EDMBs in Zintl phases, intermetallic compounds, and cluster compounds of main-group elements (with homonuclear bonds) is consistent with the well-known brittle and shiny metallic properties attributed to them.^{91,92,117} Note that EDMBs are directional bonds, due to the existence of partially localized electrons,^{1,2} which leads to brittle behavior. At the same time, EDMBs have partially delocalized electrons,^{1,2} which results in a shiny metallic aspect and moderate electrical conductivity, as expected for incipient metals (assumed to feature the new

metivalent bond by Wuttig and collaborators).^{13–15,25,26} Additionally, it must be also noticed that the formation of EDMBs in electron-rich elements should be no surprise since ERMBs can also be found in electron-deficient elements; *e.g.* hydrogen can form a 3c–4e ERMB with F atoms in the $[\text{HF}]^-$ or HF_2^- anion (cataloged as a hydrogen bond by IUPAC).^{19,50,55,118} as we have already stated in Sections 2 and 3.

5. TlTe: simultaneous presence of ERMBs and EDMBs at RP

Further examples of molecules and solids with electron-rich elements and EDMBs will be provided in future works; however, regarding the different geometries and properties of ERMBs and EDMBs, we consider important to comment here a last and paradigmatic example that will help us to distinguish between both kinds of multicenter bonds. Noteworthy, the crystalline structure of solid TlTe at room conditions exhibits simultaneously both ERMBs and EDMBs.

First of all, it must be mentioned that the structure of this solid was mentioned by Böttcher⁹¹ and by Papoian and Hoffmann,²⁰ but no satisfactory explanation for its crystalline structure was given to our knowledge. According to the most recent work of Papoian and Hoffmann, the structure can be understood on the light of the Zintl–Klemm concept by considering that each Tl atom donates one electron to each Te atom so the Te atoms become pseudo-I atoms, thus leading to the formation of homonuclear ERMBs between the Te atoms.²⁰ In contrast, we are going to show that, despite we agree that this is a Zintl compound in which the Tl atom donates charge to the Te atoms, the interpretation of all the Te–Te bonds as ERMBs is not correct. A hint that goes against the conclusion of Papoian and Hoffmann is that, despite its simple formula unit, the crystalline structure of TlTe is rather complex since it features one independent Tl atom and three independent Te atoms, so it is expected that the three different Te atoms have different atomic charges which could lead to the formation of different types of chemical bonds.

The tetragonal crystalline structure of TlTe¹¹⁹ shown in Fig. 9 has four independent atoms at the following Wyckoff sites: Tl(16k), Te₁(8h), Te₂(4b), and Te₃(4a). All Te atoms can be considered to be linked forming linear homonuclear Te–Te–Te bonds; however, we are going to show that these bonds are different (EDMBs and ERMBs) depending on the type of Te and the different electronic charges. Our *ab initio* simulations yield the following Bader [Löwdin] charges: Tl^{0.32+[0.09+]}, Te₁^{0.41-[0.14-]}, Te₂^{0.05-[0.06-]}, and Te₃^{0.39-[0.11-]}. As observed, there is a considerable difference between the Bader and Löwdin charges; however, the Bader and Löwdin charges of the Tl and Te atoms do fit with the multiplicities of all atoms. As we have done for paratellurite, the charge of Te atoms must be divided by the appropriate multiplicity ratio between the Te atom multiplicity and the multiplicity of Tl atom (16) to get the charge of Tl. For instance, the Bader charge of Tl (+0.32) can be obtained as the sum of all the renormalized Bader charges of



the Te atoms ($(-0.41/2) + (-0.05/4) + (-0.39/4) = -0.315$) if we consider that the dividing factors of the Bader charges of the Te atoms are the multiplicity ratios ($16/8 = 2$, $16/4 = 4$, and $16/4 = 4$, respectively) to be applied to the Te atoms (at 8h, 4b, and 4a Wyckoff sites, respectively).

According to our calculations, bonds between Te_3 atoms are the easiest to understand since these atoms show twofold coordination by forming infinite linear tellurium chains along the *c*-axis. The $\text{Te}_3\text{--Te}_3$ bond distance is 3.09 Å (simulated value of 3.05 Å) and its calculated ES (ET) value is 1.04 [0.65] (0 [0]). Therefore, the large value of this $\text{Te}_3\text{--Te}_3$ bond distance, when compared to the more covalent Te–Te bond distance of 2.84 Å in the A8 phase of Te at RP, and the small ES value (close to 1.0) indicates that the $\text{Te}_3\text{--Te}_3$ bonds of the infinite linear tellurium chain are E MDBs (see $\text{Te}_3\text{--Te}_3$ bonds in Fig. 1 and Fig. S1, ESI†). These E MDBs are similar to those found in the infinite linear iodine chain (Fig. 4c). This can be understood if we consider that Te_3 atoms (with a Bader charge of -0.39) behave as pseudo-I atoms, as initially assumed by Papoian and Hoffmann,²⁰ since they accept the charge donated by Tl atoms. However, they do not form ERMBs, unlike it was previously assumed,²⁰ but E MDBs, like in the infinite linear iodine chain.^{24,27}

A different case is that of fourfold coordinated Te_2 atoms. Our calculations show that they behave as pure Te atoms because they do not accept electronic charge from the Tl atoms (Bader charge of -0.05), unlike the assumption of Papoian and Hoffmann. Te_2 atoms show two types of bonds. On the one hand, Te_2 atoms also form infinite linear tellurium chains

along the *c*-axis, like Te_3 atoms, and with the same $\text{Te}_2\text{--Te}_2$ bond distance as $\text{Te}_3\text{--Te}_3$ bonds. The ES (ET) value for the $\text{Te}_2\text{--Te}_2$ bond is 1.04 [0.65] (0 [0]), much like the $\text{Te}_3\text{--Te}_3$ bond. Therefore, the $\text{Te}_2\text{--Te}_2$ bonds can also be classified as E MDBs, like the $\text{Te}_3\text{--Te}_3$ bonds. On the other hand, Te_2 atoms also form two bonds with adjacent Te_1 atoms that are perpendicular to the infinite linear tellurium chain. Therefore, the infinite linear Te_2 chain along the *c*-axis can be considered as a crosswise stacking of linear triatomic $\text{Te}_1\text{--Te}_2\text{--Te}_1$ fragments. The experimental (simulated) $\text{Te}_1\text{--Te}_2$ bond distance is 3.01 Å (2.99 Å) and exhibits a calculated ES (ET) value of 1.4 [0.88] (0.36 [0.08]). Note that the ET value is calculated as the difference between the electronic charges of the Te_2 and Te_1 atoms. Due to the large $\text{Te}_1\text{--Te}_2$ bond distance and the larger values of ES and ET than the $\text{Te}_3\text{--Te}_3$ and $\text{Te}_2\text{--Te}_2$ bonds, the $\text{Te}_1\text{--Te}_2$ bond within the linear $\text{Te}_1\text{--Te}_2\text{--Te}_1$ trimer can be classified as an ERMB, indeed the linear triatomic $\text{Te}_1\text{--Te}_2\text{--Te}_1$ fragment is a 3c–4e bond (see $\text{Te}_1\text{--Te}_2$ bond in Fig. 1). This result is consistent with the different Bader charges of the central and external atoms of this linear trimer and the accumulation of the electronic charge in the external atoms of the trimer (see Fig. 9).

We must note that our calculations of Löwdin charges and ES values with LOBSTER lead to a classification of the $\text{Te}_1\text{--Te}_2\text{--Te}_1$ bond as an E MDB (see Fig. S1, ESI†). We think that the wrong classification of this bond by LOBSTER comes from the small Löwdin charge (~ -0.14) attributed to the Te_1 and Te_3 atoms, which is related to the small Löwdin charge ($\sim +0.09$) attributed to the Tl atom. The anomalously small positive charge of Tl atoms given by LOBSTER is in contradiction with

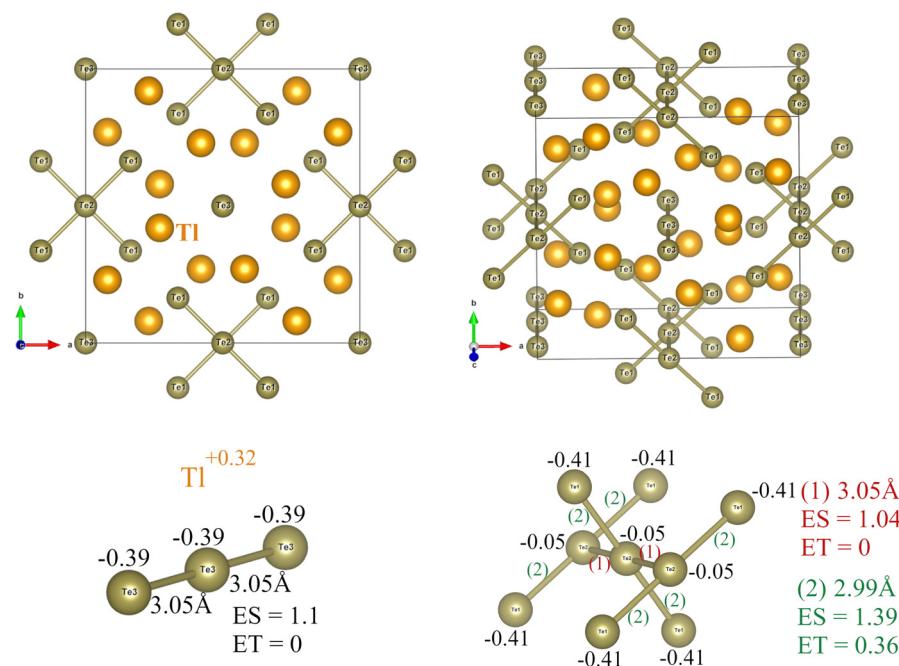


Fig. 9 Details of the crystalline structure of TlTe. Te_1 , Te_2 , and Te_3 atoms correspond to $\text{Te}(8\text{h})$, $\text{Te}(4\text{b})$, and $\text{Te}(4\text{a})$ atoms, respectively. $\text{Te}(4\text{a})$ and $\text{Te}(8\text{h})$ atoms form E MDBs and ERMBs, respectively, with the linear geometry of Fig. 10a. $\text{Te}(4\text{d})$ atoms form ERMBs and E MDBs in perpendicular directions that correspond to the square planar geometry of Fig. 10b. Details on bond distances, Bader atomic charges, ET, and ES of various bonds are illustrated here and summarized in Table S2 of the ESI†.



the original view that TlTe is a Zintl phase in which Tl atoms give their electrons to Te atoms in order to form the homonuclear bonds experimentally found in the structure and that is consistent with the already discussed pseudo-I character of Te_3 atoms.

We have demonstrated the case of TlTe, where density-based calculations show better agreement with experimental data than orbital-based calculations. This example complements previous cases, such as $\text{Sc}_2\text{Si}_2\text{O}_7$, where the opposite trend is observed. The reasons behind why some examples are better described by density-based methods while others align more closely with orbital-based methods remain unclear. These differences may stem from the inherent limitations of these methods in capturing the complexity of the multicenter bonds studied here. Nevertheless, we can confidently conclude that, in general, both density-based and orbital-based methods exhibit consistent trends (see Fig. 1 and Fig. S1, ESI[†]), underscoring the robustness of our results regardless of the method employed.

In brief, the tetragonal structure of TlTe at RP exhibits two types of Te–Te bonds: ERMBs and EDMBs. While EDMBs extend along more than three bonds forming infinite linear chains of Te_3 and Te_2 atoms along the *c*-axis, ERMBs only extend along three centers forming finite linear triatomic $\text{Te}_1\text{–Te}_2\text{–Te}_1$ fragments or trimers perpendicular to the *c*-axis. Both EDMBs and ERMBs in TlTe are longer than covalent Te–Te atoms, thus explaining the confusion in distinguishing between both types of multicenter bonds. However, ERMBs and EDMBs can be distinguished because ERMBs are finite (we will show in the next section that all pure ERMBs are indeed 3c–4e bonds), have external atoms bearing much larger electronic charge than the central atom, and the ES (also ET) values of ERMBs are usually larger than those of EDMBs and comparable or even larger than those of ionic-covalent bonds. Moreover, ERMBs are usually slightly shorter than EDMBs. In TlTe, Te–Te EDMBs are of the order of 3.05 Å, while ERMBs are of the order of 2.99 Å. The larger bond length of EDMBs than of ERMBs is consistent with the smaller ES value for EDMBs than for ERMBs and accounts for the smaller bond order of EDMBs than of ERMBs. In other words, the smaller the number of electrons shared between two atoms is, the longer the bond and the smaller the bond order, as expected from Pauling's formula for bond order.²⁴

Due to the different ES and ET values in ERMBs and EDMBs, both multicenter bonds are expected to exhibit different properties. In this context, we can briefly comment that at relatively low pressures AX_3 compounds (A = Na, K, Rb, Cs; X = Cl, Br, I) crystallize in phases showing finite linear X–X–X ERMBs, as we have shown for CsI_3 at RP in the previous section; however, AX_3 compounds undergo a phase transition at HP to a phase with infinite linear X–X–X bonds as those of Fig. 4f. In fact, a recent report has highlighted the different nature and properties of the two different types of bonds in AX_3 compounds.¹²⁰ These different properties can be understood if the infinite linear atomic bonds (typically found in the $\text{Pm}\bar{3}n$ phase) of AX_3 compounds are EDMBs as we have already suggested in this work.¹²⁰ More detailed work on these compounds to show the

pressure-induced EDMB formation from original ERMBs is in progress.¹²¹

6. A change of paradigms

The above examples have led us to propose that the paradigm that EDMBs can be found only in electron-deficient elements (not in electron-rich elements) is not valid anymore. Moreover, we have shown examples in which EDMBs are 3c–2e bonds, such as B_2H_6 and $\text{Sc}_2\text{Si}_2\text{O}_7$, and examples in which EDMBs are infinitely extended 2c–1e bonds. In addition, we have found an extensive number of examples among molecules and solids that justify that ERMBs are found only in three-center molecules in different directions, such as in 1D (I_3^- , $[\text{FHF}]^-$, XeF_2 , and I_3^- in CsI_3), in 2D (XeF_4 and Te_5^{2-} in Cs_2Te_5), and in 3D (SF_6 , XeF_6 , and TeI_6^{2-} in Cs_2TeI_6). Besides, we have reasoned that ERMBs do not show a flagrant violation of the doublet/octet rule. Under this light, we tentatively propose four new paradigms:

(1) EDMBs and ERMBs can be found in molecules and solids made of electron-rich elements.

(2) EDMBs, present in either electron-rich or electron-deficient elements, are found as 3c–2e bonds or as infinitely extended 2c–1e bonds.

(3) Pure ERMBs can be formed only in linear or quasi-linear three-center molecules. This means that all ERMBs are 3c–4e bonds extended either in one, two, or three dimensions. Corollary: multicenter bonds extended to more than three centers in one dimension, either in a linear or a zigzag way, cannot be pure ERMBs and tend to form EDMBs.

(4) EDMBs and ERMBs do not violate, in general, the doublet/octet rule.

Since the first two paradigms are clear from the examples analyzed in ref. 1 and 2 and in the previous sections, we are going to provide in this section additional arguments to support the third paradigm. We will show that the lack of linear ERMBs longer than three centers is supported by symmetry and energy arguments and that angular or zigzag ERMBs are not pure ERMBs. The fourth paradigm will be discussed in Section 8.

6.1. Energy arguments

Lubchenko and coworkers have tentatively proposed a rule about a two-center localized molecular orbital that states:³⁹ “*in a stable (or metastable) molecule, the bonding localized molecular orbitals should be two-center and cover each nearest neighbor bond. Conversely, when the molecule is not fully covered by two-center localized molecular orbitals, the molecule is subject to structural instability*”. In other words, a system with n centers and $(n + 1)$ bonding electrons is stable if there are enough electrons to fill $(n - 1)$ localized molecular orbitals. This can be translated into the equation $n + 1 = 2(n - 1)$, whose only solution is $n = 3$.³⁹

This result shows that the 3c–4e ($n = 3$) bond is the only stable molecule with n centers and $n + 1$ bonding electrons and



allows explaining the large number of linear three-center molecules with four electrons experimentally found in 1D, 2D, and 3D,³⁹ and commented in the previous sections. It must be stated that the energy argument agrees with simulations that prove that linear polyiodide anions with n atoms and $n + 1$ bonding electrons that are longer than three centers; *i.e.*, longer than I_3^- , have different bond lengths along the chain and are either directly unstable or weak and easily perturbed.^{24,122}

6.2. Symmetry arguments

According to Lubchenko and coworkers,³⁹ multicenter bonds become electron deficient for any linear molecule longer than three centers. We have recently extended this reasoning to linear and angular polyiodides^{24,27} thanks to *ab initio* calculations.¹²² Moreover, the geometry of different polyiodides has been recently discussed. The preference of angular geometries for polyiodides longer than three centers has been justified as a way to reinforce the chains when weak EDMBs occur in linear polyiodides longer than three centers.²⁴

As an example, we show in Fig. 4c–f the possible polymerization of the H_2 and I_2 molecules in infinite linear and zigzag chains. The infinite linear iodine chain in Fig. 4f is a clear example of EDMB (extended 2c–1e bond) similar to the infinite linear hydrogen chain in Fig. 4e. The only difference between the two infinite linear chains is that, unlike in the infinite linear hydrogen chain, the I atoms in the infinite linear iodine chain each I atom exhibits six non-bonding electrons, corresponding to three LEPs, distributed around each I atom in the plane perpendicular to the 2c–1e bonds, which are in a linear p-type configuration. Note that the six electrons of the LEPs in the infinite linear iodine chain can be thought to form a sp^2 configuration in the plane perpendicular to the 2c–1e bonds thus resulting in a total bipyramidal geometry (symmetry requirements show that the six electrons are distributed forming a toroid when ELF isosurfaces are plotted (see toroidal ELF in HF in Fig. S5a, ESI[†]).

The bonding in the infinite linear chains of H and I atoms is different from that in the infinite zigzag chain of H and F atoms (Fig. 4c) and the infinite zigzag iodine chain (Fig. 4d), as already commented in ref. 24 and 27. The infinite zigzag chains show ERMBs because terminal atoms of the linear branches do not have six non-bonding electrons but only four; *i.e.*, two LEPs instead of three. This means that terminal halogen atoms have three bonding electrons that contribute to the formation of 3c–4e bonds in both HF and I zigzag chains. As already commented, these zigzag chains can be understood as a concatenation of 3c–4e ERMBs.^{24,27} This result agrees with the view of Dronskowski and coworkers for the infinite zigzag iodine chain.⁸⁵ Therefore, we conclude that the different symmetry of the two infinite atomic chains in Fig. 4c–f results in a different bond type. While the infinite linear atomic chain bears EDMBs, the infinite zigzag chain bears ERMBs.

The above examples prove that the comparison of bonding in PCMs, like β -GeTe, and in polyiodides (in general) done in ref. 85 is not valid since the type of bonding in polyiodides depends on the geometry and size of the polyiodide chain.^{24,27}

Noteworthy, the infinite linear iodine chains have been experimentally observed in AX_3 compounds ($A = Na, K, Rb, Cs; X = Cl, Br, I$).¹²⁰ These halides have the same kind of bonds that occur in the nominally assumed Sb^{2-} atoms in Li_2Sb and Te_3^- atoms in $TlTe$, as already discussed. This is not surprising since these charged atoms can be considered pseudo-I atoms within the Zintl–Klemm perspective. Contrarily, the ERMBs shown in Fig. 4d for the infinite zigzag iodine chain have been observed in iodine at HP^{85} and also occur in Te^- atoms (again pseudo-I atoms) in UTe_5 .⁹¹

It is important to stress that the ERMBs occurring in the infinite zigzag atomic chains that result from the coalescence or concatenation of 3c–4e bonds are not pure ERMBs, unlike those of the isolated molecules with linear 3c–4e bonds. This conclusion is supported by our calculations. For instance, the ES value of the F–H bond in the F–H–F 3c–4e bond of the HF_2^- molecule is much larger than that of the F–H bond in the F–H–F bond in the *Cmcm* HP phase of HF at 20 GPa (*i.e.*, when solid HF exhibits infinite zigzag chains of H and F atoms as that shown in Fig. 4c), despite both bonds can be classified as ERMBs. The reason is that the concatenation of 3c–4e bonds even in a zigzag form leads to an increase in the bond length of the two-center bonds within ERMB (with respect to the pure 3c–4e bond) and to a loss of shared electrons in line with Pauling's formula for bond order as recently justified for polyiodides.⁶⁴ Consequently, ERMBs in infinite zigzag atomic chains are not pure ERMBs and can be considered as weakened ERMBs in the sense that they tend to become EDMBs as the length of the bonds increases with the increase of atoms in the chain beyond three centers.

A possible explanation for the decrease of the ES value in concatenated 3c–4e bonds in HF with respect to pure isolated 3c–4e bonds in HF^{2-} could be due to the reduction in the number of electrons available for bonding in solid HF with respect to HF^{2-} . This reduction arises from the stoichiometry of the system: in solid HF there is a 1:1 ratio of H and F atoms that has fewer electrons for bonding compared to isolated HF^{2-} ions with a 1:2 stoichiometry. In both cases, four electrons are expected to participate in the bonding of three atoms and the rest of electrons are considered to be part of LEPs; however, the smaller number of electrons in LEPs in HF than in HF^{2-} ions, could lead to a smaller ES value in HF than in HF^{2-} if those extra electrons participate with a small contribution to the bonding. This argument is the same as the electron count hypothesis previously commented in Section 4.2.1.

Other examples of materials in which symmetry plays a role are crystalline solids, *e.g.* $BaZnSb_2$ and sc-Po, since the translational symmetry is the main one in crystalline solids. As already mentioned, all Sb atoms of the planar array in $BaZnSb_2$ are at 4c sites, so they are all equivalent and must have the same Bader [Löwdin] charge (-0.6 [-0.82]). Similarly, all Po atoms in sc-Po are located in the 4a Wyckoff position, so all atoms are equivalent and must have the same charge as already found in ref. 1 and 2. In both compounds, Sb–Sb and Po–Po bonds have $ET = 0$, so ERMBs cannot be found simply because ERMBs are characterized by relatively high ET values. Therefore,



symmetry prevents ERMBs from occurring in homonuclear bonds of atoms occupying the same Wyckoff position in solids. In other words, we have already reasoned that the central atom of a trimer has always a more positive charge than the external atoms within a 3c–4e bond due to the different Wyckoff sites for the central and terminal atoms of the trimer. Consequently, the translational symmetry in solids makes it impossible for the central atom of a given or chosen trimer to shift electronic charge to the terminal parts of the unit provided that all atoms of the trimer have the same Wyckoff site. This symmetry argument makes it impossible to establish ERMBs in an infinite linear atomic chain with all atoms having the same Wyckoff site, as in the infinite linear chain of H and I atoms.

A good example of the impossibility of ERMBs to be formed in infinite linear molecules, even when these molecules have atoms at different Wyckoff sites with different electronic charges, is the crystalline β phase of PCMs; *e.g.* IV–VI chalcogenides, such as β -GeTe (Fig. 8c). In β -GeTe, Ge and Te are located at 4a and 4b sites and have Bader [Löwdin] charges of +0.36 [+0.18] and –0.36 [–0.18], respectively. Since Ge and Te occupy different Wyckoff sites in β -GeTe, the impossibility commented in the above paragraph does not apply and we could *a priori* assume that a 3c–4e ERMB could be possible along three perpendicular axes around the Ge atom with positive Bader [Löwdin] charge and a $\text{Te}^{\delta-}\text{Ge}^{\delta+}\text{Te}^{\delta-}$ configuration (according to our previous reasoning of (– + –) charge distribution in the I_3^- molecule). However, a 3c–4e ERMB cannot be possible around the Te atom with negative Bader [Löwdin] charge and $\text{Ge}^{\delta+}\text{Te}^{\delta-}\text{Ge}^{\delta+}$ configuration. Note that this configuration would result in a (+ – +) charge distribution that has not been observed in any 3c–4e molecule yet. The reason is that this charge configuration will concentrate the charge at the central atom of the trimer, and this will lead to a flagrant violation of the octet rule; a situation that tends to be avoided in Nature,⁵⁷ and that does not occur in 3c–4e bonds as we have shown in Fig. 5. On the other hand, no five-center bond with $\text{Te}^{\delta-}\text{Ge}^{\delta+}\text{Te}^{\delta-}\text{Ge}^{\delta+}\text{Te}^{\delta-}$ configuration or larger ERMBs with equal Ge–Te bond distances could be possible, as already discussed for polyiodides.²⁴ A similar linear five-atom configuration is observed in the linear I_5^- molecule which is only slightly stable¹²² and is not observed isolated but only in confined spaces like in nanotubes.⁷⁹ It must be stressed that the linear I_5^- molecule does not show ERMBs as recently discussed,^{24,122} since this molecule can be understood as a system composed of two external I_2 molecules weakly linked to a central I^- anion, as shown by their different I–I bond distances. Only angular V- or L-shape I_5^- molecules (*i.e.*, showing a zigzag geometry), with ERMBs in a configuration similar to that of Fig. 4c and d, are stable as already discussed.^{24,122}

Regarding the possibility of linear molecules with ERMBs larger than three centers (4c–6e, 5c–6e, 5c–8e, *etc.*), we have to stress that they have been thoroughly studied,^{54,85,123} but it must be clarified that these linear molecules do not have the equal or nearly equal bond distances as those observed in 3c–4e ERMBs, as already discussed for the linear I_4^- , I_4^{2-} , and I_5^- polyanions, and in EDMBs (either in 3c–2e bonds or in general

in extended 2c–1e bonds).^{24,122} In any case, we want to stress that the issue is different in molecules, even inside solids, where several symmetry restrictions must not be obeyed.

It must be noted that there are ways to overcome the limitation of the extension of ERMBs to three centers and extend ERMBs to infinite in one or more dimensions. One example of the infinite extension of ERMBs in 2D is found in Cs_2Te_5 (Fig. 6b). In this solid, terminal Te_2 atoms in planar Te_5^{2-} molecular units form short pure covalent $\text{Te}_2\text{–Te}_2$ bonds that are almost perpendicular to the planar Te_5^{2-} molecular units with $\text{Te}_2\text{–Te}_1\text{–Te}_2$ ERMBs. The covalent bonds link the planar Te_5^{2-} molecular units in two directions.¹²⁴ In other words, since ERMBs cannot be extended beyond three centers, the alternation of covalent bonds and ERMBs in a zigzag way is a way to extend ERMBs in different directions. A similar example is the recent finding of Ru–Ru–Ru trimers featuring 3c–4e bonds in the low-temperature monoclinic phase of crystalline RuP.¹²⁵ In this solid, the central and external Ru atoms of the trimer occupy different 4e Wyckoff sites so the Ru–Ru–Ru trimers can be extended along one direction forming zigzag-like ladders because the external Ru atoms form covalent-like Ru–Ru bonds almost perpendicular to the direction of the 3c–4e bonds. These two examples are similar to the polymerized infinite zigzag chains schematized in Fig. 4c and d that also show linked 3c–4e bonds due to the change of direction to avoid the previously mentioned restrictions for the formation of ERMBs. In summary, we consider that ERMBs with equal bond distances are not possible in linear molecules of more than three atoms inside crystalline solids. Consequently, infinite linear multicenter bonds in crystalline solids cannot be ERMBs and must be necessarily EDMBs.

As a final comment regarding the difference between ERMBs and EDMBs, we want to comment that it could be argued against our claim for the formation of linear EDMBs in the octahedral phases of pnictogens and chalcogens and crystalline PCMs² that linear 3c–2e EDMBs are rare in molecules,¹²⁶ whereas linear 3c–4e ERMBs are common in molecules.⁵⁴ In this regard, it must be emphasized that two types of 3c–2e bonds are known: supported and unsupported 3c–2e bonds.¹⁰ It has been commented that the 3c–2e bonds of H_3^+ and B_2H_6 are of the supported type and tend to be bent, while the B–H–B and Al–H–Al 3c–2e bonds in B_2H_7^- and Al_2H_7^- molecules are of the unsupported type.¹⁰ It has been proposed that these unsupported 3c–2e bonds can be either bent or linear: bent in the absence of a crystal lattice and linear when the crystal lattice is present; *i.e.*, when a translational symmetry is imposed.^{127,128} Therefore, the existence of linear unsupported EDMBs in crystalline solids can also be justified due to the presence of the translational symmetry of the crystal lattice. This argument is in line with the symmetry arguments mentioned in this section.

Moreover, in this comparison between ERMBs and EDMBs, we consider that there is an anti-symmetry between ERMBs and EDMBs. The infinite extension of ERMBs along one direction, which occurs thanks to the formation of infinite zigzag chains as a concatenation of 3c–4e bonds, has its anti-symmetrical



counterpart in EDMBs since the infinite extension of EDMBs along one direction occurs thanks to the formation of infinite linear chains with 2c-1e bonds, which can be considered a concatenation of 3c-2e bonds. Curiously, 3c-4e bonds, which are linear or quasi-linear, become extended by forming infinite zigzag chains, whereas 3c-2e bonds, which are typically non-linear, become extended by forming infinite linear chains. Therefore, the 2c-1e EDMBs in the infinite linear atomic chains in the octahedrally-coordinated phases of pnictogens, chalcogens, PCMs, and in the twofold-coordinated phases of H (Fig. 4c) and halogens (Fig. 4d)²⁴ can be considered as the equivalent to the weakened ERMBs in the infinite zigzag atomic chains in HX compounds (X = F, Cl, Br, I) and in polyhalogens.²⁴ This concept will be developed in more detail in a forthcoming article.

7. Hypercoordinated units in ERMBs and EDMBs

In this section, we want to make a pertinent observation about the hypercoordinated multicenter units with different geometries that have been experimentally observed upon the formation of both ERMBs and EDMBs in electron-rich elements; units whose geometries go beyond those of the classical geometries of iono-covalent bonds.

As already stated, electron-rich elements participating in hypercoordinated multicenter units have been traditionally associated with the presence of ERMBs or hypervalent 3c-4e bonds. In the literature, the concept of hypervalent molecules of main-group elements has been related with the violation of the $8 - N$ rule for atomic coordination, where N is the total number of valence s and p electrons and $8 - N$ indicates the number of covalent 2c-2e bonds to be formed to complete the octet.¹²⁹⁻¹³¹ In other words, hypervalent molecules show elements with a larger atomic coordination (hypercoordination) than expected if all bonds were considered single covalent 2c-2e bonds.

Frequently, the violation of the $8 - N$ rule has been interpreted (see ref. 60 and 132) as if these hypercoordinated molecules, with assumed 2c-2e bonds, would have also violated the Lewis-Langmuir doublet/octet rule,^{133,134} i.e., it has been considered that the central atom of a 3c-4e bond is surrounded by more than the two (eight) atoms allowed for s (s + p) orbitals. However, as we have already discussed regarding Fig. 5, several works have suggested that hypervalent molecules, such as XeF_2 or I_3^- , do not violate the doublet/octet rule.^{57,58,135-141} Although there is still a doubt whether there is a violation of the octet rule for hypervalent molecules of the second kind (PF_5 and SF_6) and not for those of the first kind (XeF_2 or I_3^-).^{142,143} Moreover, it has been proposed that the terms "hypervalence" and "hypervalent" should be sent to the graveyard and replaced by more convenient terms, like "hypercoordinated" or "hypo-bound".^{57,60,135-141}

Recent works by Grabowski on molecules have discussed the formation of hypervalent units driven by various non-covalent interactions in secondary bonds, including hydrogen, triel,

tetrel, pnictogen, chalcogen, and halogen bonds. Grabowski has pointed out that the mechanism of formation of these secondary bonds is the same in all of them and related to the σ -hole model.^{58,144} Moreover, it has been suggested that the hypervalent units formed by secondary bonds, leading to 3c-4e ERMBs in molecules follow the rules of the VSEPR model.^{49,81,132}

Unfortunately, the linear ERMBs have not been clearly distinguished in the diagrams of the hypervalent units in previous works in our opinion,^{49,58,132,144} so we have plotted in Fig. 10 some of the simplest hypercoordinated units with ERMBs according to the VSEPR model⁸¹ with a special emphasis on the location of the linear multicenter bonds (green bonds). As a matter of example of ERMBs, the XeF_2 molecule, and I_3^- ion in CsI_3 exhibit the linear geometry of Fig. 10a, the XeF_4 molecule shows the square planar geometry of Fig. 10b, and the TeI_6 molecule in Cs_2TeI_6 features the cubic geometry of Fig. 10c.

The replot of the simplest hypercoordinated units with linear ERMBs of Fig. 10 is also interesting for EDMBs since we have shown in this work and in previous ones^{1,2} that those hypercoordinated multicenter units are also observed in EDMBs in which electron-rich elements participate. Notice that the linear three-center geometries in Fig. 10 (in most cases related to the presence of LEPs in electron-rich elements) are in general different from those found in EDMBs of electron-deficient elements, which have no LEPs.¹²⁶ Notably, the 3D EDMBs of the A_h phase of α -Po show the cubic geometry of Fig. 10c and the A_i phase of β -Po shows the rhombohedral geometry of Fig. 10c.^{1,2} Similarly, the cubic geometry of Fig. 10c has been observed in 3D EDMBs of PCMs with rs structure, such as $\beta\text{-GeTe}$, SnTe , PbS , PbSe , and PbTe , and the rhombohedral geometry of Fig. 10c has been observed in 3D EDMBs of PCMs with tetradymite-like structure, such as $\alpha\text{-Bi}_2\text{Se}_3$, $\beta\text{-As}_2\text{Te}_3$, $\alpha\text{-Sb}_2\text{Te}_3$, and $\alpha\text{-Bi}_2\text{Te}_3$. On the other hand, the 1D Sb-Sb (Te-Te) EDMBs of the infinite linear atomic chains in Li_2Sb (TlTe) show the linear geometry of Fig. 10a and the 1D Te-O EDMBs and the covalent bonds of TeO_2 show the see-saw geometry of Fig. 10a. The 2D Sb-Sb EDMBs of the planar Sb array in BaZnSb_2 (Fig. 7b) and the mixture of 2D Te-Te ERMBs and EDMBs of Te(4b) in TlTe exhibit the square planar geometry of Fig. 10b, while the 2D Sn-Se EDMBs and the covalent bond in *Cmcm*-type SnSe (Fig. 8b) exhibit the square pyramidal geometry of Fig. 10b. This last reasoning also applies to isostructural InBr, InI, and TlI.

In view of these results, we conclude that the hypercoordinated multicenter units shown in Fig. 10 represent some of the simplest linear multicenter bonds that comply with VSEPR rules, regardless of whether they are ERMBs (3c-4e) or EDMBs (3c-2e or extended 2c-1e). In the hypercoordinated units, multicenter bonds can be directed in 1D, 2D, and 3D, alone or in combination with LEPs and iono-covalent bonds, which are noted as E and C, respectively in Fig. 10. Therefore, our work shows that linear hypercoordinated multicenter units such as those shown in Fig. 10 occur for both ERMBs and EDMBs in molecules and solids made of electron-rich elements.



Hypercoordinated Multicenter Units

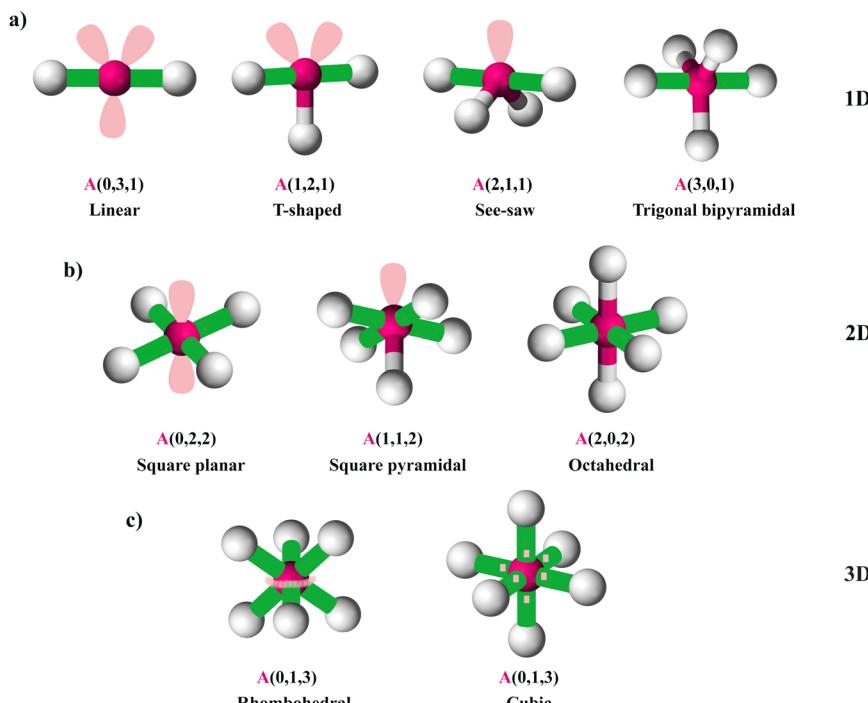


Fig. 10 Hypercoordinated multicenter units around a central electron-rich A atom (pink color) showing the simplest linear three-center bonds in (a) one (1D), (b) two (2D), and (c) three (3D) dimensions. Each unit around the A atom is defined by three numbers enclosed in parentheses, $A(C, E, M)$, denoting the number of ionic-covalent bonds (C), LEPs (E), and linear multicenter bonds (M) that are depicted in white-pink, light pink, and green color, respectively. These units are typical of molecular units with electron-rich multicenter bonds (ERMBs) that specifically correspond to 3c–4e bonds. These units have been observed in molecules, such as I_3^- , XeF_2 , XeF_4 , and SF_6 , and solids, such as CsI_3 , Cs_2Te_5 , $TlTe$, and Cs_2Tel_6 . However, these 1D, 2D, and 3D units extended indefinitely also appear in solids with electron-deficient multicenter bonds (EDMBs), such as the A_h and A_i phases of group-15 and -16 elements,² in IV–VI and V_2 – VI_3 PCMs, such as β -GeTe and Bi_2Te_3 , and in other solids, e.g. Li_2Sb , $TlTe$, and $BaZnSb_2$. In (c), the light pink halo (squares) indicates the presence of a stereochemically weakly active (inactive) LEP in the rhombohedral (cubic) geometries, as it happens in the A_i (A_h) phases of group-16 (group-15) elements (see ref. 2).

We must note in passing that no “hypervalency” occurs in multicenter bonds. It was assumed that the atomic hypercoordination was caused because the atom was acting with a valence higher than expected (hence the term “hypervalency”);¹⁷ e.g. Ge and Te atoms in β -GeTe have 2^+ and 2^- valence, respectively, but both atoms exhibit a sixfold coordination (hypocoordination) as if they would act with a valence 6^+ and 6^- , respectively. However, there is no hypervalency since the hypercoordinated atoms do not show a valency beyond their maximum allowed values. Note that the atom valency is defined as the capacity of the atom to share one electron in each formed covalent bond; since the bonds in hypercoordinated units are not all covalent bonds, but there are multicenter bonds, the assumed hypervalency is not correct. Consequently, we suggest, in agreement with previous works,^{57,60,135–141} that the terms “hypervalency” and “hypervalent” should be replaced by “hypocoordination” and “hypercoordinated”, as shown in Fig. 10, because atomic hypocoordination means the presence of multicenter bonds in the involved atom. We note that the terms “hypobonding” or “hypobonded” can complement “hypocoordination” and “hypercoordinated” when describing multicenter bonds. In this context, an atom is said to form a

“hypobond” when the bond involves fewer electrons shared than in a pure covalent bond, as is characteristic of multicenter bonding.

It must be stressed that the formation of hypercoordinated multicenter units corresponding to 3c–4e ERMBs is clearly illustrated in Fig. 10, as ERMBs cannot extend beyond three centers. However, the formation of hypercoordinated multicenter units corresponding to 2c–1e E MDBs that are extended beyond three centers (along 1D, 2D, or 3D) in solids could lead to misinterpretation of Fig. 3d–f and 10 since in these figures the bonds are represented only for three-center molecules. In this context, it must be understood that the formation of extended 2c–1e E MDBs (along 1D, 2D, or 3D) is similar to the mechanism shown in Fig. 3d–f and the geometries of Fig. 10 but extended infinitely in the corresponding directions. For instance, in As-I phase (A7 structure) the primary covalent As–As bond is almost linearly linked by secondary bonds to two As atoms of neighbor layers (As–As–As–As) along one direction;² thus, the *trans* influence of the two As–As secondary bonds on the primary covalent As–As bond is on both sides of the central As₂ molecule. Moreover, since each As atom in the A7 phase of As participates in three, almost perpendicular



primary As–As bonds, this picture should be extended in the three, almost perpendicular directions. In this way, the A_h structure of the As-II(sc-As) phase at HP is formed by extended 3D EDMBs² (not by 3c–4e bonds as previously assumed²⁰), unlike what Fig. 3d–f and 10 could initially suggest due to their limited schematic view. The polymerization reactions schematized in Fig. 4e and f help us to illustrate the extension of EDMBs shown in Fig. 3d–f and 10 in one direction.

Now that we have demonstrated that the geometries of hypercoordinated units present in Fig. 10 can be found in both ERMBs and EDMBs in electron-rich elements, we propose a new notation for these hypercoordinated multicenter units. Our notation is different from that used by Crabtree⁵⁰ and Grabowski⁵⁸ who use the number of ligands, L, which can be understood as the number of bonds, but do not distinguish between ERMBs and EDMBs (note that previous diagrams considered only ERMBs⁵⁸). Here we propose that the hypercoordinated multicenter units of Fig. 10 might be noted as $A(C, E, M)$, where A refers to the symbol of the hypercoordinated central atom A, while the numbers C, E, and M correspond to the number of ionic-covalent bonds, LEPs, and linear multicenter bonds (both ERMBs or EDMBs), respectively.

To illustrate the application of our notation with various compounds exhibiting EDMBs, we can cite: (i) the $A(0,3,1)$ unit of chains of Sb atoms in Li_2Sb ⁹⁶ and of Te(4a) atoms in $TlTe$;¹⁰³ (ii) the $A(1,2,1)$ unit of O atom in $SbOF$ (see Fig. S8a of the ESI[†]);¹⁴⁵ (iii) the $A(2,1,1)$ unit of Te atoms in TeO_2 ¹⁰³ and of Sb atoms in $SbOF$;¹⁴⁵ (iii) the $A(3,0,1)$ units of P atoms in $TiPO_4$ –V at 48 GPa (see Fig. S8b of the ESI[†]);¹⁴⁶ (iv) the square planar $A(0,2,2)$ unit of Sb atoms at the planar array in $BaZnSb_2$ or of Te(4b) in $TlTe$;¹⁰³ (v) the square pyramidal $A(1,1,2)$ units present around Sn (Tl) and Se (I) atoms in *Cmcm*-type $SnSe$ (TII);¹⁰² and (vi) the rhombohedral and cubic $A(0,1,3)$ units of the A_h and A_i phases in group-15 and -16 elements as well as of many PCMs of IV–VI and V₂–VI₃ families with rs and tetradymite-like structures, respectively,² including crystalline GST, the reference material for phase change electronic storage.

As we have mentioned, the geometries in Fig. 10 are different from those of EDMBs found in electron-deficient elements; however, we would like to add that a similar geometry to that of the linear geometry in Fig. 10a would be found for the case of the infinite linear hydrogen bond (see Fig. 4e). This could be considered a $A(0,0,1)$ unit (not shown in Fig. 10 but is similar to the $A(0,3,1)$ unit without the three LEPs). The $A(0,0,1)$ unit of hydrogen could also correspond to the type of bond in the atomic/polymeric hydrogen at HP,⁷⁰ in which H–H bonds form a ring instead of an infinite linear atomic chain.

8. Is the doublet/octet rule violated in the hypercoordinated units of ERMBs and EDMBs?

In Fig. 5 we have already discussed that the central atom in ERMBs, such as those in HF_2^- , XeF_2 , and I_3^- , does not flagrantly violate the doublet/octet rule; however, we would like

to conclude this paper with a broader discussion on whether the central A atom in the hypercoordinated multicenter $A(C, E, M)$ units of Fig. 10 violates or not the doublet/octet rule when M is either an ERMB or an EMB.

To verify whether the doublet/octet rule is satisfied or not, we can consider, as a first approximation, that: (i) each ionic-covalent bond contains two electrons (the bonding electron pair); (ii) each LEP contains two electrons (the non-bonding electron pair); and (iii) each multicenter bond is characterized by partially sharing two electrons between every two atoms (ERMB) or only one electron between every two atoms (EMB). In this way, we can easily check that, in general, the $A(C, E, M)$ units of Fig. 10 for EDMBs seem to verify the octet rule, although some cases need further discussion. The verification of the octet rule is not so clear for ERMBs,⁶⁰ despite Grabowski has reasoned that the formation of the hypercoordinated units in molecules with 3c–4e ERMBs is connected with the mechanism of the σ -hole model to uphold the doublet rule for alkali and alkaline earth metals and the octet rule in main-group elements.^{57,58} Let us comment here in more detail on the possible violation of the doublet/octet rule in both ERMBs and EDMBs.

Regarding ERMBs, we have already commented in Section 2, when talking about the electron distribution in hypercoordinated XeF_2 , XeF_4 , and XeF_6 molecules, that the octet rule is not severely violated, unlike proposed in ref. 60, since it can be considered that this violation of the octet rule occurs in a second approximation and not in a first approximation. The reason is that it can be considered that there are only 8 electrons in the valence electron sphere close to the central Xe atom and 2, 4, and 6 additional electrons in the van der Waals electron sphere (at longer distance from the central Xe atom) in XeF_2 , XeF_4 , and XeF_6 , respectively.⁶⁰ In any case, there is still a doubt whether there is a violation of the octet rule for hypervalent molecules of the second kind (PF_5 and SF_6) and not for those of the first kind (XeF_2 or I_3^-).^{142,143} In this regard, we consider that all pure ERMBs (either of first or second kind) are 3c–4e bonds, so the reasoning of Fig. 5 regarding I_3^- can be applied to every pure ERMB in 1D, 2D or 3D to verify that the doublet/octet rule is in general satisfied as a first approximation.

Regarding EDMBs, it has been suggested that the octet rule is being satisfied in molecules⁶² and the same is expected in solids. Taking into account the above considerations, eight electrons seem to surround each A atom in hypercoordinated units with only one EMB (1D geometry), such as in the $A(0,3,1)$, $A(1,2,1)$, $A(2,1,1)$, and $A(3,0,1)$ units of Fig. 10a. The same occurs for the units with two EDMBs (2D geometry), the $A(0,2,2)$, $A(1,1,2)$, and $A(2,0,2)$ units of Fig. 10b, and for the units with three EDMBs (3D geometry), such as the cubic $A(0,1,3)$ units of Fig. 10c.

A clear example of the validity of the octet rule for EDMBs with the geometry of the $A(0,1,3)$ unit is the A_h phase of pnictogens, *e.g.* the sc-As phase.² Since pnictogens have two s-type and three p-type valence electrons, the A_h phase features two s-type electrons forming part of the inactive LEP (they are



distributed into six lobes, as shown by pink regions in the cubic $A(0,1,3)$ unit of Fig. 10c), while the three p-type electrons participate in the three mutually perpendicular EDMBs. Therefore, the central A atom of the cubic $A(0,1,3)$ units has eight valence electrons, thus satisfying the octet rule. It must be emphasized that this result is contrary to the previous assumption that hypercoordinated units in the crystalline phases of PCMs, such as β -GeTe (isoelectronic to sc-As), led to a violation of the octet rule.¹³² The same reasoning would apply to consider the validity of the doublet rule for H, when it forms EDMBs at HP. In particular, polymeric H ($\cdots H-H-H \cdots$) either in ring form at HP⁷⁰ or in the infinite linear chain of Fig. 4e could be considered to form hypercoordinated $A(0,0,1)$ units (with no LEPs) with only one EDMB (1D geometry). Therefore, two electrons surround each H atom, thus satisfying the doublet rule. It can be observed that the EDMBs of I atoms in the infinite linear iodine chain (Fig. 4f) lead I atoms to obey the octet rule.

In summary, assuming that each ionic-covalent bond, each LEP, and each EDMB, accounts for two electrons in multicenter hypercoordinated $A(C, E, M)$ units, the doublet and octet rules are satisfied for the central A atom if we consider the equation $2(C + E + M) = 2$ or 8 . For the octet rule, it has to be satisfied that $C + E + M = 4$. This is exactly the condition satisfied by all the $A(C, E, M)$ units of Fig. 10. For the doublet rule (case of H), it has to be satisfied that $C + M = 1$ since H has no LEP ($E = 0$). In this context, H has only three bonding possibilities to satisfy the doublet rule, either it has a covalent bond ($C = 1, M = 0$), as in H_2 , an ERMB ($C = 0, M = 1$), as in HF_2^- , or an EDMB ($C = 0, M = 1$), as it is expected to occur in atomic/polymeric hydrogen at HP.

Interestingly, the octet rule seems to be violated in the hypercoordinated units with EDMBs for group-16 elements, *e.g.* the cubic and rhombohedral $A(0,1,3)$ units of α -Po and β -Po.^{1,2} In this case, in addition to the six p-type electrons of the three EDMBs around each central A atom (three coming from the central A atom), one has to count the three additional electrons of the A atom (two s-type electrons corresponding to the LEP plus the extra p electron in chalcogens compared to pnictogens). This extra electron is distributed either among the six lobes of the cubic $A(0,1,3)$ unit (see light pink cubes in Fig. 10c) typical of α -Po or in the toroidal halo of the rhombohedral $A(0,1,3)$ unit (see light pink monosynaptic toroidal in Fig. 10c) typical of β -Po. Therefore, a total number of nine electrons seems to be around each hypercoordinated EDMB unit for octahedrally-coordinated chalcogens, thus violating the octet rule.

Perhaps the violation of the octet rule in the hypercoordinated (octahedrally coordinated) units with EDMBs for group-16 elements in the A_h and A_i phases is only apparent. A study of the prevalence of the cubic *vs.* the rhombohedral units in AF_6E molecules with ERMBs, where A is the central atom, F is the ligand atom, and E is the LEP, has concluded that both the cubic O_h symmetry, like that of the cubic unit of Fig. 10c, or the C_{3v} symmetry, like that of the rhombohedral unit of Fig. 10c, in AF_6E molecules, such as XeF_6 , are very close in energy.¹⁴¹

The predominance of one geometry over the other seems to be related to a very fine balance between ligand-ligand repulsions and the energy gained by the expansion of the two non-bonding electrons of the LEP in the valence shell. This balance has been suggested to mainly depend on the atomic radii ratio between the central atom and the ligands. Interestingly, the small energy difference between the two configurations for AF_6E molecules with ERMBs is similar to the one we have found between the two polymorphs of Po at 0 K, which explains that β -Po tends to α -Po above 2 GPa.² With the above considerations, it could be speculated that, similarly to the case of AF_6E molecules with ERMBs,¹⁴⁷ the cubic and rhombohedral configurations of the two phases of Po at RP likely occur because of the fine balance of the two aforementioned energy terms (which certainly will depend on the temperature and pressure conditions).

Regarding the violation of the octet rule, it has also been suggested in the same work of AF_6E molecules with ERMBs, that the non-bonding electrons (being part of LEPs) in the AF_6E molecules could behave as a mixture of valence-core electrons so on average they contribute with only one valence electron; *i.e.*, one of the two electrons of the s-type LEP could be considered part of the core.¹⁴⁷ Therefore, it could also be speculated that the three non-bonding electrons of chalcogens in the A_h and A_i phases likely behave as in AF_6E molecules with ERMBs, so that on average they contribute with only two valence electrons, that summed to the other six p-type electrons of the three EDMBs would make the central atom to satisfy the octet rule. This hypothesis of the mixture of valence-core electrons for s-type electrons can be justified in the A_h and A_i phases of chalcogens due to the large energy difference between s and p states in chalcogens.¹⁴⁸ In summary, this reasoning could be a possible explanation for the fulfilment of the octet rule in both the cubic and rhombohedral geometries with EDMBs in the A_h and A_i phases of chalcogens. Moreover, this reasoning could provide an answer to the question, posed by Papoian and Hoffmann²⁰ and already commented in our previous work,² about how could it be possible the existence of the cubic α -Po phase taking into account the electron counting rule suggested by Papoian and Hoffmann for the formation of multicenter bonds.²⁰ Note that we have already commented that the electron counting rule of Papoain and Hoffmann for the formation of multicenter bonds is correct.² However, they applied it incorrectly to explain the bonds in many Zintl phases. They interpreted all multicenter bonds in materials of electron-rich elements as ERMBs; a feature that we have shown in this work and in ref. 2 to be incorrect because some of them are EDMBs.

To close this section, it is important to stress that the already commented rocksalt structure of β -GeTe at HP, the orthorhombic $Cmcm$ structure of SnSe at HP and TII at RP, and the paratellurite (α -TeO₂) structure at RP, would constitute examples of the violation of the octet rule if we discard the presence of EDMBs, as it has been repeatedly suggested.²⁰ The rocksalt structure of β -GeTe shows Ge and Te atoms with sixfold coordination. This means that if all were single covalent bonds,



a total of 12 valence electrons (6 bonds \times 2 electrons per bond) would lead to a flagrant violation of the octet rule in addition to hypervalence, as already discussed. If we assume that they are ERMBs, as Hoffmann and his collaborators suggest,^{20,85} they would also violate the octet rule since there are *ca.* 1.5 electrons shared in any ERMB (6 bonds \times 1.5 electrons per bond = 9 electrons). Only considering that they are EDMBs one gets a total of 8 valence electrons around Ge and Te atoms as it should be to complete the octet (6 bonds \times 1 electrons per bond = 6 electrons plus the two electrons from the inactive s-type LEP). Similarly, the *Cmcm* structure of SnSe at HP with fivefold coordination for Sn and Se atoms cannot be understood without the participation of EDMBs. Note that if all bonds were single covalent bonds a total 12 valence electrons would be around each Sn and Se atom (5 bonds \times 2 electrons per bond = 10 electrons plus 2 electrons from the LEP which accounts for the layered structure of SnSe). Therefore, both Sn and Se atoms would strongly violate the octet rule. The same would occur if the four bonds of SnSe in the layer plane were considered ERMBs since 10 electrons would be around each Sn and Se atom (4 bonds \times 1.5 electrons per bond = 6 electrons plus two electrons from the covalent bond and two electrons from the LEP). On the other hand, metallic bonds cannot be invoked to explain the structure and properties of β -GeTe and SnSe at HP. In particular, the layered *Cmcm* structure of SnSe and the relatively strong directionality of the bonds within the layers are incompatible with the presence of non-directed metallic bonds but are compatible with the presence of directed EDMBs. The mystery for SnSe can be solved if we consider that each atom has 4 valence electrons in the direction perpendicular to the layers (coming from one iono-covalent bond and one LEP) plus 4 valence electrons in the layer plane (corresponding to each of the 4 2c-1e bonds that form the 2D EDMBs).

Similar arguments can be invoked to explain the structure and properties of paratellurite, in which Te atoms are fourfold coordinated to O atoms. If the four bonds were single covalent, there would be 4 bonds \times 2 electrons per bond = 8 electrons, which added to the two electrons from the LEP present in Te that result in a trigonal bipyramidal geometry, leads to a total of 10 valence electrons around Te atoms. The mystery here is solved if we consider that two collinear Te-O bonds are EDMBs and provide 2 electrons in total and not 4, thus resulting in the 8 valence electrons around the Te atoms. Therefore, our results explain why there are two short and two long bonds in α -TeO₂ without the need to invoke the existence of a thermal variation in bond lengths.¹⁴⁹ In summary, we conclude that the doublet/octet rule is not violated, in general, in multicenter bonds. Moreover, the doublet/octet rule and the VSEPR theory are fundamental tools to realize that both ERMBs and EDMBs can be naturally integrated in a unified theory of multicenter bonding.

9. Conclusions

ERMBs and EDMBs are two types of multicenter bonds that are relatively well known in molecules but remain poorly studied in

solids. We have shown that both ERMBs and EDMBs can occur in molecules and solids made of electron-rich elements. In our recent work,² it has been shown that the octahedrally-coordinated phases of α -Po and β -Po in pnictogens and chalcogens, respectively, and typically found at HP, are characterized by EDMBs; a result that can be extrapolated to the crystalline phases of PCMs, such as β -GeTe. Such a result is in contrast to the previous metavalent bond model for PCMs, which considers that PCMs feature 2c-1e bonds, and to the previous hypervalent bond model for PCMs, which considers that crystalline PCMs are characterized by ERMBs (also known as hypervalent bonds or hyperbonds).

In this work, which is a continuation of ref. 2, we have addressed the nature of EMB and ERMBs in several molecules and solids, most of them having electron-rich elements. We have shown that ERMBs and EDMBs have the same origin; *i.e.*, they come from a mixture of primary iono-covalent bonds and secondary non-covalent bonds, and a similar formation mechanism, which consists of three stages; however, they differ in the electronic charge reorganization at stage 2. The non-bonding electrons of the stereochemically active LEP become bonding electrons when the ERMB is formed thus providing the two electrons needed for the new bond. Contrarily, the non-bonding electrons of the stereochemically active LEP tend to remain as non-bonding delocalized electrons in a weakly stereochemically active or even inactive LEP when the EMB is formed and the charge needed for the new bond mainly comes from the primary iono-covalent bond.

Both ERMBs and EDMBs are usually longer (by more than 0.2 Å) than the covalent bond. We have clarified that quantum mechanical calculations, leading to the 2D ES *vs.* ET map, can help distinguishing both EDMBs and ERMBs since the two kinds of multicenter bonds are located at different positions with respect to the classical ionic, covalent, and metallic bonds.

The ERMB is a multicenter bond that can be considered the extension of the single polar covalent (2c-2e) bond to three centers, giving rise to the well-known 3c-4e bond. We have shown that the ERMB has less than two electrons shared between two atoms, but it is not a 2c-1e bond, like the EMB. In other words, the EMB has approximately half the electronic charge of a pure single covalent (2c-2e) bond. The simplest EMB is the 3c-2e bond typical of boranes, although it can be also found as extended 2c-1e bonds in solids.

We have provided several examples of electron-rich systems that exhibit ERMBs, all of which involve 3c-4e bonds. We have also shown that several compounds made of electron-rich elements, such as Li₂Sb, BaZnSb₂, and α -Po, feature indeed homopolar EDMBs in 1D, 2D, and 3D, respectively, despite they were previously assumed to exhibit ERMBs.²⁰ Also examples of heteropolar EDMBs in β -GeTe, SnSe, TeO₂, ScSi₂O₇, and SbPO₄ as well as a paradigmatic example with simultaneous homopolar ERMBs and EDMBs (TlTe) have been discussed.

We have shown that all EDMBs fall in the same region of the ES *vs.* ET map than the central B-H bonds of the B₂H₆ molecule. We have reasoned that the bonds found in atomic/polymeric hydrogen at HP must be EDMBs since the mechanism



of strengthening of the secondary, intermolecular bonds is the same for hydrogen, B_2H_6 , pnictogens, chalcogens, halides, and crystalline PCMs, such as $\beta\text{-GeTe}$. All these materials have been shown to feature extended EDMBs of unsupported type at certain pressures.

We have provided the simplest geometries of linear or quasi-linear hypercoordinated multicenter units (both with ERMBs and EDMBs) around a central electron-rich A element. These units do not obey the $8 - N$ rule, have geometries compatible with the VSEPR model, and are not hypervalent despite being formed by electron-rich elements. We suggest that the “hypervalent” units should be renamed as “hypercoordinated multicenter units”. Besides, we have proposed a new notation (A(C,E,M)) to classify hypercoordinated units that include multicenter bonds which is based on the number of ionic-covalent bonds, LEPs, and multicenter bonds (either ERMBs or EDMBs).

We have justified that in extended solids hypercoordination of electron-rich elements with linear or quasi-linear bonds is only possible for EDMBs since ERMBs can only appear in three-center molecular units (even in solids) as 3c-4e bonds due to energy and symmetry restrictions. ERMBs cannot form linear bonds longer than three centers without severely violating the octet rule for the internal atoms of the molecule. Extension of ERMBs along one direction usually needs to alternate ERMBs (alone or in combination with other types of bonds) in a zigzag configuration due to the impossibility of doing it in a linear configuration. This limitation does not hold for 3c-2e EDMBs, which extend to infinite as colinear 2c-1e bonds.

Finally, we have reasoned about the violation of the doublet/octet rule for the central A atom in hypercoordinated multicenter units (mainly those formed by electron-rich elements). We have shown that, in general, the doublet/octet rule is satisfied around the central A atom for hypercoordinated units with EDMBs. For hypercoordinated units with ERMBs, the central A atom seems to obey the octet rule as a first approximation and to violate the octet rule only in a second approximation. This is justified by the distribution of atomic electrons in concentric electronic (core, valence, and van der Waals) spheres.

All the results summarized here have allowed us to tentatively propose four new paradigms:

(1) EDMBs and ERMBs can be found in molecules and solids made of electron-rich elements.

(2) EDMBs, present in either electron-rich or electron-deficient elements, are found as 3c-2e bonds or as infinitely extended 2c-1e bonds.

(3) Pure ERMBs can be formed only in linear or quasi-linear three-center molecules. This means that all ERMBs are 3c-4e bonds extended either in one, two, or three dimensions. Corollary: multicenter bonds extended to more than three centers in one dimension, either in a linear or a zigzag way, cannot be ERMBs and tend to form EDMBs.

(4) EDMBs and ERMBs do not violate, in general, the doublet/octet rule.

In summary, we have presented in this work a novel unified theory of multicenter bonding that is consistent with the VSEPR

rules and the doublet/octet rules and that has challenged several paradigms related to the ERMBs and EDMBs. For this reason, we consider that the results of this work have very far-reaching consequences for the broad scientific community, not only for supramolecular chemists but also for condensed matter scientists, since the clear distinction between ERMBs and EDMBs we have provided here will allow explaining the structures (and also the properties) found in many materials from simple elements to complex materials, such as PCMs, highly efficient thermoelectrics, topological insulators, superconductors, highly efficient photovoltaic materials, Zintl phases, intermetallics, and cluster compounds at RP, as well as in atomic/polymeric nitrogen and hydrogen at HP. This understanding will open the door for a better understanding of the chemical bonding mechanisms in the above-commented advanced materials for improving their performance. In summary, we hope that this work will promote further work to understand ERMBs and EDMBs in solids and their associated exceptional property portfolio.

Data availability

The data supporting this article has been included as part of the ESI.[†]

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This publication is financed by Spanish Ministerio de Ciencia e Innovación and the Agencia Estatal de Investigación MCIN/AEI/10.13039/501100011033 as part of the project MALTA Consolider Team network (RED2022-134388-T), and I + D + i projects PID2022-138076NB-C42/C44 co-financed by EU FEDER funds, by project PROMETEO CIPROM/2021/075 (GREENMAT) financed by Generalitat Valenciana and co-financed by EU FEDER. This study also forms part of the Advanced Materials programme supported by MCIN with funding from European Union NextGenerationEU (PRTR-C17.I1) and by Generalitat Valenciana through project MFA/2022/025 (ARCANGEL). We would like to express our gratitude to A. Otero-de-la-Roza, J. Contreras-García, Á. Lobato, F. Izquierdo, M. Savastano, and Á. Vegas for their insightful and engaging discussions on our findings. Their input and feedback greatly enriched the development of this article. More specifically, A. Otero-de-la-Roza has assisted us with the CRITIC2 code; Á. Lobato and J. Contreras-García called our attention to the parallelism between the mechanism of EMB formation and the bond formation in the atomic/polymeric phases of nitrogen and hydrogen, respectively; Á. Lobato and F. Izquierdo are acknowledged by collaborating in the study of multicenter bonds in AX_3 compounds; Á. Vegas draw our attention to various systems, including $\text{Sc}_2\text{Si}_2\text{O}_7$; and M. Savastano is acknowledged for his collaboration on the study of



polyiodides and for suggesting us the works of J. Echeverría and S. Alvarez.

References

- H. H. Osman, A. Otero-de-la-Roza, P. Rodríguez-Hernández, A. Muñoz and F. J. Manjón, DOI: [10.26434/chemrxiv-2023-pv66p-v2](https://doi.org/10.26434/chemrxiv-2023-pv66p-v2).
- H. H. Osman, A. Otero-de-la-Roza, P. Rodríguez-Hernández, A. Muñoz and F. J. Manjón, *J. Mater. Chem. C*, 2024, **12**, 10447–10474.
- R. E. Rundle, *J. Am. Chem. Soc.*, 1947, **69**, 1327–1331.
- R. E. Rundle, *J. Chem. Phys.*, 1949, **17**, 671–675.
- R. E. Rundle, *J. Phys. Chem.*, 1957, **61**, 45–50.
- H. C. Longuet-Higgins, *Quarterly Reviews*, Chemical Society, 1957, **11**, 121–133.
- L. Pauling, *The Nature of the Chemical Bond and the Structure of Molecules and Crystals: an Introduction to Modern Structural Chemistry*, Cornell Univ. Press, Ithaca, New York, 1960.
- K. Wade, *Electron deficient compounds*, Nelson, London, 1971.
- C. Zheng and R. Hoffmann, *Z. Naturforsch. B*, 1986, **41**, 292–320.
- R. L. DeKock and W. B. Bosma, *J. Chem. Educ.*, 1988, **65**, 194.
- W. M. Lipscomb, The Boranes and their Relatives.
- R. F. W. Bader, *Atoms in molecules: a quantum theory*, Clarendon Press, Oxford, 1990.
- J.-Y. Raty, M. Schumacher, P. Golub, V. L. Deringer, C. Gatti and M. Wuttig, *Adv. Mater.*, 2019, **31**, 1806280.
- Y. Cheng, O. Cojocaru-Mirédin, J. Keutgen, Y. Yu, M. Küpers, M. Schumacher, P. Golub, J.-Y. Raty, R. Dronskowski and M. Wuttig, *Adv. Mater.*, 2019, **31**, 1904316.
- M. Wuttig, C. Schön, D. Kim, P. Golub, C. Gatti, J. Raty, B. J. Kooi, Á. M. Pendás, R. Arora and U. Waghmare, *Adv. Sci.*, 2023, **35**, 2308578.
- P. C. Müller, S. R. Elliott, R. Dronskowski and R. O. Jones, *J. Phys.: Condens. Matter*, 2024, **36**, 325706.
- J. I. Musher, *Angew. Chem., Int. Ed. Engl.*, 1969, **8**, 54–68.
- R. J. Hach and R. E. Rundle, *J. Am. Chem. Soc.*, 1951, **73**, 4321–4324.
- G. C. Pimentel, *J. Chem. Phys.*, 1951, **19**, 446–448.
- G. A. Papoian and R. Hoffmann, *Angew. Chem., Int. Ed.*, 2000, **39**, 2408–2448.
- B. Silvi, *J. Mol. Struct.*, 2002, **614**, 3–10.
- R. Ponec, G. Yuzhakov and D. L. Cooper, *Theor. Chem. Acc.*, 2004, **112**, 419–430.
- I. Mayer, *J. Mol. Struct.:THEOCHEM*, 1989, **186**, 43–52.
- M. Savastano, H. H. Osman, Á. Vegas and F. J. Manjón, *Chem. Commun.*, 2024, **60**, 12677–12689.
- M. Wuttig, C.-F. Schön, J. Lötfering, P. Golub, C. Gatti and J.-Y. Raty, *Adv. Mater.*, 2023, **35**, 2208485.
- M. Wuttig, V. L. Deringer, X. Gonze, C. Bichara and J.-Y. Raty, *Adv. Mater.*, 2018, **30**, 1803777.
- F. J. Manjón, H. H. Osman, M. Savastano and Á. Vegas, *Materials*, 2024, **17**, 2840.
- A. Otero-de-la-Roza, E. R. Johnson and V. Luña, *Comput. Phys. Commun.*, 2014, **185**, 1007–1018.
- S. Maintz, V. L. Deringer, A. L. Tchougréeff and R. Dronskowski, *J. Comput. Chem.*, 2016, **37**, 1030–1035.
- M. Giambiagi, M. S. de Giambiagi and K. C. Mundim, *Struct. Chem.*, 1990, **1**, 423–427.
- M. S. De Giambiagi, M. Giambiagi and M. De Souza Fortes, *J. Mol. Struct.:THEOCHEM*, 1997, **391**, 141–150.
- R. Ponec and I. Mayer, *J. Phys. Chem. A*, 1997, **101**, 1738–1741.
- P. Bultinck, R. Ponec and S. Van Damme, *J. Phys. Org. Chem.*, 2005, **18**, 706–718.
- E. Matito, *Phys. Chem. Chem. Phys.*, 2016, **18**, 11839–11846.
- F. Feixas, M. Solà, J. M. Barroso, J. M. Ugalde and E. Matito, *J. Chem. Theory Comput.*, 2014, **10**, 3055–3065.
- P. C. Müller, C. Ertural, J. Hempelmann and R. Dronskowski, *J. Phys. Chem. C*, 2021, **125**, 7959–7970.
- J. Hempelmann, P. C. Müller, C. Ertural and R. Dronskowski, *Angew. Chem., Int. Ed.*, 2022, **61**, e202115778.
- G. A. Landrum and R. Hoffmann, *Angew. Chem., Int. Ed.*, 1998, **37**, 1887–1890.
- J. C. Golden, V. Ho and V. Lubchenko, *J. Chem. Phys.*, 2017, **146**, 174502.
- M. Savastano, *Dalton Trans.*, 2021, **50**, 1142–1165.
- R. C. L. Money-Slater, *Acta Crystallogr.*, 1959, **12**, 187–196.
- W. Grochala, R. Hoffmann, J. Feng and N. W. Ashcroft, *Angew. Chem., Int. Ed.*, 2007, **46**, 3620–3642.
- E. Espinosa, I. Alkorta, J. Elguero and E. Molins, *J. Chem. Phys.*, 2002, **117**, 5529–5542.
- R. M. Lobayan, R. C. Bochicchio and A. D. Marturet, *J. Mol. Graphics Model.*, 2022, **112**, 108121.
- T. H. Lee and S. R. Elliott, *Adv. Mater.*, 2020, **32**, 2000340.
- J.-P. Gaspard, *Phys. Status Solidi RRL*, 2022, **16**, 2200111.
- J. Echeverría and S. Alvarez, *Chem. Sci.*, 2023, **14**, 11647–11688.
- J. Echeverría and S. Alvarez, *Chem. Sci.*, 2024, **15**, 12166–12168.
- T. H. Lee and S. R. Elliott, *Phys. Status Solidi RRL*, 2021, **15**, 2000516.
- R. H. Crabtree, *Chem. Soc. Rev.*, 2017, **46**, 1720–1729.
- H. A. Bent, *Chem. Rev.*, 1968, **68**, 587–648.
- T. Clark, M. Hennemann, J. S. Murray and P. Politzer, *J. Mol. Model.*, 2007, **13**, 291–296.
- J. S. Murray, P. Lane, T. Clark and P. Politzer, *J. Mol. Model.*, 2007, **13**, 1033–1038.
- R. D. Harcourt, *Bonding in Electron-Rich Molecules*, Springer International Publishing, 2016.
- G. A. Landrum, N. Goldberg and R. Hoffmann, *J. Chem. Soc., Dalton Trans.*, 1997, 3605.
- N. W. Alcock, in *Advances in Inorganic Chemistry and Radiochemistry*, ed. H. J. Emeléus and A. G. Sharpe, Academic Press, 1972, vol. 15, pp. 1–58.
- S. J. Grabowski, *Phys. Chem. Chem. Phys.*, 2017, **19**, 29742–29759.
- S. J. Grabowski, *Molecules*, 2021, **26**, 4939.

59 G. Parkin, *J. Chem. Educ.*, 2023, **100**, 4644–4652.

60 S. Noury, B. Silvi and R. J. Gillespie, *Inorg. Chem.*, 2002, **41**, 2164–2172.

61 K. S. Pitzer, *J. Am. Chem. Soc.*, 1945, **67**, 1126–1132.

62 G. Parkin, *J. Chem. Educ.*, 2019, **96**, 2467–2475.

63 G. R. Eaton, *Found. Chem.*, 2023, **25**, 285–298.

64 A. D. Walsh, *J. Chem. Soc.*, 1947, 89–92.

65 T. Shimajiri, S. Kawaguchi, T. Suzuki and Y. Ishigaki, *Nature*, 2024, **634**, 347–351.

66 M. Wuttig, C.-F. Schön, M. Schumacher, J. Robertson, P. Golub, E. Bousquet, C. Gatti and J.-Y. Raty, *Adv. Funct. Mater.*, 2022, **32**, 2110166.

67 V. Labet, P. Gonzalez-Morelos, R. Hoffmann and N. W. Ashcroft, *J. Chem. Phys.*, 2012, **136**, 074501.

68 V. Labet, R. Hoffmann and N. W. Ashcroft, *J. Chem. Phys.*, 2012, **136**, 074502.

69 V. Labet, R. Hoffmann and N. W. Ashcroft, *J. Chem. Phys.*, 2012, **136**, 074503.

70 V. Riffet, V. Labet and J. Contreras-García, *Phys. Chem. Chem. Phys.*, 2017, **19**, 26381–26395.

71 H. Nakatsuji, *Acc. Chem. Res.*, 2012, **45**, 1480–1490.

72 P. G. Johannsen, W. Helle and W. B. Holzapfel, *J. Phys. Colloq.*, 1984, **45**, C8-199–C8-201.

73 A. Ikram, B. H. Torrie and B. M. Powell, *Mol. Phys.*, 1993, **79**, 1037–1049.

74 T. Kume, T. Tsuji, S. Sasaki and H. Shimizu, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1998, **58**, 8149–8151.

75 E. Katoh, H. Yamawaki, H. Fujihisa, M. Sakashita and K. Aoki, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1999, **59**, 11244–11250.

76 D. Duan, F. Tian, Z. He, X. Meng, L. Wang, C. Chen, X. Zhao, B. Liu and T. Cui, *J. Chem. Phys.*, 2010, **133**, 074509.

77 L. Zhang, Y. Wang, X. Zhang and Y. Ma, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2010, **82**, 014108.

78 H. Mao and R. J. Hemley, *Rev. Mod. Phys.*, 1994, **66**, 671–692.

79 A. A. Tonkikh, D. V. Rybkovskiy and E. D. Obraztsova, *J. Phys. Chem. C*, 2023, **127**, 3005–3012.

80 J. Contreras-García, Á. M. Pendás, B. Silvi and J. M. Recio, *J. Phys. Chem. B*, 2009, **113**(4), 1068–1073.

81 R. J. Gillespie, *Coord. Chem. Rev.*, 2008, **252**, 1315–1327.

82 R. Hoffmann, *Solids and surfaces: A chemist's view of bonding in extended structures*, Wiley-VCH, 1988.

83 E. Wigner and H. B. Huntington, *J. Chem. Phys.*, 1935, **3**, 764–770.

84 R. Hoffmann, (Cornell University, Ithaca, NY, USA). *Private communication*.

85 J. Hempelmann, P. C. Müller, L. Reitz and R. Dronkowski, *Inorg. Chem.*, 2023, **62**, 20162–20171.

86 P. Schneiderhan, P. Schmidt, M. Ströbele, C. P. Romao and H.-J. Meyer, *Cryst. Growth Des.*, 2020, **20**, 3780–3784.

87 S. Wei, J. Wang, S. Deng, S. Zhang and Q. Li, *Sci. Rep.*, 2015, **5**, 14393.

88 I. Vázquez-Fernández, S. Mariotti, O. S. Hutter, M. Birkett, T. D. Veal, T. D. C. Hobson, L. J. Phillips, L. Danos, P. K. Nayak, H. J. Snaith, W. Xie, M. P. Sherburne, M. Asta and K. Durose, *Chem. Mater.*, 2020, **32**, 6676–6684.

89 A. Ruth, K. Németh, K. C. Harkay, J. Z. Terlik, L. Spentzouris and J. Terry, *J. Appl. Phys.*, 2013, **113**, 183703.

90 T. Poreba, S. Racioppi, G. Garbarino, W. Morgenroth and M. Mezouar, *Inorg. Chem.*, 2022, **61**, 10977–10985.

91 P. Böttcher, *Angew. Chem., Int. Ed. Engl.*, 1988, **27**, 759–772.

92 Á. Vegas, *Structural Models of Inorganic Crystals*, Universitat Politècnica de València, 1st edn, 2018.

93 D. Akhtar, V. D. Vankar, K. L. Goel and T. C. Chopra, *J. Mater. Sci.*, 1979, **14**, 988–994.

94 M. Brylak, M. H. Möller and W. Jeitschko, *J. Solid State Chem.*, 1995, **115**, 305–308.

95 R. Nesper, *Prog. Solid State Chem.*, 1990, **20**, 1–45.

96 R. Gérardin and J. Aubry, *C. R. Acad. Sci., Ser. C: Sci. Chim.*, 1974, **278**, 1097–1098.

97 M. Brylak and W. Jeitschko, *Z. Naturforsch., B*, 1994, **49**, 747–752.

98 Y. Wang, K. Wang, Y. Ma, M. Zhou, H. Wang and G. Liu, *J. Phys.: Condens. Matter*, 2021, **33**, 355403.

99 I. Loa, R. J. Husband, R. A. Downie, S. R. Popuri and J.-W. G. Bos, *J. Phys.: Condens. Matter*, 2015, **27**, 072202.

100 M. Xu, S. Jakobs, R. Mazzarello, J.-Y. Cho, Z. Yang, H. Hollermann, D. Shang, X. Miao, Z. Yu, L. Wang and M. Wuttig, *J. Phys. Chem. C*, 2017, **121**, 25447–25454.

101 I. Efthimiopoulos, M. Berg, A. Bande, L. Puskar, E. Ritter, W. Xu, A. Marcelli, M. Ortolani, M. Harms, J. Müller, S. Speziale, M. Koch-Müller, Y. Liu, L.-D. Zhao and U. Schade, *Phys. Chem. Chem. Phys.*, 2019, **21**, 8663–8678.

102 L. Helmholz, *Z. Kristallogr. – Cryst. Mater.*, 1936, **95**, 129–137.

103 J. Leciejewicz, *Z. Kristallogr.*, 1961, **116**, 345–353.

104 F. Folger, *Z. Anorg. Allg. Chem.*, 1975, **411**, 111–117.

105 K. Kohn, K. Inoue, O. Horie and S.-I. Akimoto, *J. Solid State Chem.*, 1976, **18**, 27–37.

106 A. L. de J. Pereira, D. Santamaría-Pérez, R. Vilaplana, D. Errandonea, C. Popescu, E. L. da Silva, J. A. Sans, J. Rodríguez-Carvajal, A. Muñoz, P. Rodríguez-Hernández, A. Mujica, S. E. Radescu, A. Beltrán, A. Otero-de-la-Roza, M. Nalin, M. Mollar and F. J. Manjón, *Inorg. Chem.*, 2020, **59**, 287–307.

107 D. Errandonea, H. H. Osman, R. Turnbull, D. Diaz-Anichtchenko, A. Liang, J. Sanchez-Martin, C. Popescu, D. Jiang, H. Song, Y. Wang and F. J. Manjon, *Mater. Today Adv.*, 2024, **22**, 100495.

108 Á. Vegas, R. Notario, E. Chamorro, P. Pérez and J. F. Liebman, *Acta Crystallogr.*, 2013, **69**, 163–175.

109 R. N. Grimes, *Carboranes*, Academic Press, 3rd edn, 2016.

110 R. Llusar, A. Beltrán, J. Andrés, S. Noury and B. Silvi, *J. Comput. Chem.*, 1999, **20**, 1455–1592.

111 S. Shaik, D. Danovich, B. Silvi, D. L. Lauvergnat and P. C. Hiberty, *Chem. – Eur. J.*, 2005, **11**, 6358–6371.

112 Y. Yang, *J. Phys. Chem. A*, 2012, **116**, 10150–10159.

113 S. Shaik, D. Danovich, J. M. Galbraith, B. Braïda, W. Wu and P. C. Hiberty, *Angew. Chem., Int. Ed.*, 2020, **59**, 984–1001.

114 B. Braïda and P. C. Hiberty, *Nat. Chem.*, 2013, **5**, 417–422.

115 J. S. Murray, P. Lane, T. Clark, K. E. Riley and P. Politzer, *J. Mol. Model.*, 2012, **18**, 541–548.

116 P. Politzer and J. S. Murray, *Theor. Chem. Acc.*, 2012, **131**, 1114.

117 H. G. von Schnerring, *Angew. Chem., Int. Ed. Engl.*, 1981, **20**, 33–51.

118 S. Grabowski, *Crystals*, 2015, **6**, 3.

119 K. Stöwe, *J. Solid State Chem.*, 2000, **149**, 123–132.

120 Y. Yin, A. Aslandukova, N. Jena, F. Trybel, I. A. Abrikosov, B. Winkler, S. Khandarkhaeva, T. Fedotenko, E. Bykova, D. Laniel, M. Bykov, A. Aslandukov, F. I. Akbar, K. Glazyrin, G. Garbarino, C. Giacobbe, E. L. Bright, Z. Jia, L. Dubrovinsky and N. Dubrovinskaia, *JACS Au*, 2023, **3**, 1634–1641.

121 E. Bandiello, A. Lobato, F. Izquierdo, H. H. Osman, A. Muñoz, P. Rodríguez-Hernández, and F. J. Manjón, to be published.

122 L. J. Sæthre, O. Gropen, J. Sletten, T. Pedersen, L. B. Zinner, F. Lehrich, C. J. Nielsen, D. L. Powell and M. Trætteberg, *Acta Chem. Scand.*, 1988, **42a**, 16–26.

123 S. Hayashi, T. Nishide and W. Nakanishi, *Heteroat. Chem.*, 2018, **29**, 1–12.

124 P. Boettcher and U. Kretschmann, *Z. Anorg. Allg. Chem.*, 1982, **491**, 39–46.

125 D. Hirai, K. Kojima, N. Katayama, M. Kawamura, D. Nishio-Hamane and Z. Hiroi, *J. Am. Chem. Soc.*, 2022, **144**, 17857–17864.

126 S. P. Gnanasekar and E. Arunan, *Aust. J. Chem.*, 2020, **73**, 767–774.

127 K. Raghavachari, P. von R. Schleyer and G. W. Spitznagel, *J. Am. Chem. Soc.*, 1983, **105**, 5917–5918.

128 A. M. Sapse and L. Osorio, *Inorg. Chem.*, 1984, **23**, 627–628.

129 H. G. Grimm and A. Sommerfeld, *Z. Phys.*, 1926, **36**, 36–59.

130 W. Hume-Rothery, *London, Edinburgh Dublin Philos. Mag. J. Sci.*, 1930, **9**, 65–80.

131 W. B. Pearson, *Acta Crystallogr.*, 1964, **17**, 1–15.

132 T. H. Lee and S. R. Elliott, *Nat. Commun.*, 2022, **13**, 1458.

133 G. N. Lewis, *J. Am. Chem. Soc.*, 1916, **38**, 762–785.

134 I. Langmuir, *J. Am. Chem. Soc.*, 1919, **41**, 1543–1559.

135 P. von R. Schleyer, *Chem. Eng. News*, 1984, **62**, 4.

136 L. Suidan, J. K. Badenhoop, E. D. Glendening and F. Weinhold, *J. Chem. Educ.*, 1995, **72**, 583.

137 O. J. Curnow, *J. Chem. Educ.*, 1998, **75**, 910.

138 W. B. Jensen, *J. Chem. Educ.*, 2006, **83**, 1751.

139 J. M. Galbraith, *J. Chem. Educ.*, 2007, **84**, 783.

140 B. A. Jackson, J. Harshman and E. Miliordos, *J. Chem. Educ.*, 2020, **97**, 3638–3646.

141 N. C. Norman and P. G. Pringle, *Chemistry*, 2022, **4**, 1226–1249.

142 M. Häser, *J. Am. Chem. Soc.*, 1996, **118**, 7311–7325.

143 M. Kaupp, *The Chemical Bond*, John Wiley & Sons, Ltd, 2014, pp. 1–24.

144 S. J. Grabowski, *Science*, 2022, **4**, 17.

145 A. Åström and S. Andersson, *J. Solid State Chem.*, 1973, **6**, 191–194.

146 M. Bykov, E. Bykova, M. Hanfland, H. Liermann, R. K. Kremer, R. Glaum, L. Dubrovinsky and S. van Smaalen, *Angew. Chem., Int. Ed.*, 2016, **55**, 15053–15057.

147 J. Pilmé, E. A. Robinson and R. J. Gillespie, *Inorg. Chem.*, 2006, **45**, 6198–6204.

148 S. Deng, J. Köhler and A. Simon, *Angew. Chem., Int. Ed.*, 2006, **45**, 599–602.

149 E. R. Barney, A. C. Hannon and D. Holland, *J. Phys. Chem. C*, 2012, **116**, 3707–3718.

