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Anthraquinone-modified triazine rich g-C₃N₄ for high efficiency photocatalytic H₂O₂ synthesis via promoting singlet oxygen conversion

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Photocatalytic hydrogen peroxide (H₂O₂) production via a two-electron oxygen reduction reaction (2e⁻ ORR) represents a sustainable alternative to the energy-intensive anthraquinone (AQ) process. Although graphitic carbon nitride (g-C₃N₄) demonstrates significant advantages in photocatalytic H₂O₂ synthesis, its efficiency is severely limited by rapid charge recombination and the competitive oxidation of the critical superoxide radical ($\cdot\text{O}_2^-$) intermediate to singlet oxygen (¹O₂). Herein, AQ-modified triazine rich g-C₃N₄ (AQ/CN-x%) was successfully constructed through molten salt-assisted polycondensation and amidation grafting reactions. The coexistence of triazine and heptazine units not only promotes the separation of photogenerated charges but also provides more modification sites for AQ anchoring. Due to its strong electron-withdrawing nature, the AQ modification further enhances the separation of photogenerated charge carriers. More importantly, the AQ moiety effectively converts the competitively generated ¹O₂ into H₂O₂ via hydroanthraquinone intermediates, significantly improving the photocatalytic H₂O₂ synthesis efficiency in pure water. The optimized AQ/CN-70% catalyst achieved a remarkable H₂O₂ production rate of 165.3 μmol g⁻¹ h⁻¹, representing 4.6-fold and 13.8-fold enhancements over CN-70% and g-C₃N₄, respectively. This work provides a novel strategy for converting ¹O₂ into H₂O₂ by incorporating strongly electron-withdrawing AQ units into the triazine rich g-C₃N₄ framework, leading to a significant enhancement in photocatalytic H₂O₂ synthesis activity.

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Introduction

As a high-value multifunctional reagent, hydrogen peroxide (H₂O₂) is extensively employed in wastewater treatment, medical disinfection, and industrial bleaching.¹ Contemporary industrial H₂O₂ manufacturing predominantly employs the anthraquinone (AQ) process, which suffers from significant drawbacks, including reliance on precious-metal catalysts, substantial organic waste generation, and high energy consumption.² Photocatalytic H₂O₂ synthesis from H₂O/O₂ and solar energy has emerged as a promising sustainable alternative.³ Designing and preparing efficient and stable photocatalysts is one of the key technologies to promote the photocatalytic synthesis of H₂O₂. To date, inorganic semiconductors, metal-organic frameworks (MOFs), covalent organic frameworks (COFs) and conjugated polymers have been explored as photocatalysts to produce H₂O₂.⁴⁻⁷ Among these

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photocatalysts, graphitic carbon nitride (g-C₃N₄) is considered particularly promising for photocatalytic H₂O₂ production due to its stable physicochemical properties, ease of modification and low-cost synthesis. However, the photocatalytic efficiency for H₂O₂ generation from H₂O and O₂ remains limited due to the rapid recombination of photogenerated charge carriers and strong competition from singlet oxygen (¹O₂) generation.⁸ Unfortunately, little attention has been paid to the utilization and transformation of ¹O₂ during the photocatalytic synthesis of H₂O₂.^{9,10}

In photocatalytic processes, ¹O₂ is usually generated through two pathways: (1) energy transfer from excited dye molecules to the triplet ground state of O₂ or (2) oxidation of superoxide radicals (·O₂⁻) by photogenerated holes. These competing pathways inevitably lead to energy loss and depletion of critical intermediates, thereby reducing the efficiency of photocatalytic H₂O₂ production. Recently, it has been reported that ¹O₂ plays a crucial role in promoting photocatalytic H₂O₂ generation. For instance, the unique resonance energy transfer process from the strong excitonic effect of g-C₃N₄ can effectively activate O₂ to generate high-energy ¹O₂.¹¹ Thus, an efficient ¹O₂-engaged one-step two-electron pathway simplified the reaction procedure and optimized the intermediate energy levels, which improved the production activity of H₂O₂ by 12.2 times.¹² Through excitation energy and electron transfer, the dioxygen molecule can be converted to a highly active singlet oxygen intermediate and then is reduced to H₂O₂ *via* the photogenerated electrons with a reduced barrier.^{13,14} However, the pronounced exciton effect in g-C₃N₄ usually hinders charge carrier separation and transfer.¹⁵ Recently, a [4 + 2] cycloaddition reaction was proposed to convert ¹O₂ into H₂O₂. The generation of ¹O₂ through the oxidation of the superoxide radical (·O₂⁻), followed by its conversion into endoperoxides *via* [4 + 2] cycloaddition, facilitates charge separation and enhances the selectivity of O₂ reduction to H₂O₂.¹⁶ From the perspective of energy efficiency optimization, developing a strategy that can not only promote the separation of photo-generated charges but also rapidly convert ¹O₂ into H₂O₂ holds great promise for enhancing the efficiency of H₂O₂ photosynthesis.

Herein, AQ-modified triazine rich g-C₃N₄ (AQ/CN-*x*%) was synthesized *via* molten salt assisted polycondensation followed by an amidation grafting reaction.^{17,18} The role of AQ units in photocatalytic H₂O₂ production was systematically investigated. Combined experimental and theoretical analyses demonstrate that AQ units, acting as electron acceptors, substantially enhance charge separation efficiency. More importantly, the AQ/H₂AQ redox cycle effectively facilitates the conversion of O₂ into H₂O₂.¹⁸ This suggests that AQ may also mediate the photocatalytic transformation of ¹O₂ to H₂O₂, revealing a previously overlooked catalytic capability. The optimized AQ/CN-70% exhibited a H₂O₂ production rate of 165.3 μmol g⁻¹ h⁻¹ under ambient atmospheric conditions, representing a 13.8-fold increase over pristine g-C₃N₄. This work presents an innovative design strategy involving AQ functionalization of g-C₃N₄ to simultaneously improve charge separation efficiency and promote the conversion of ¹O₂ to H₂O₂, thereby boosting the overall efficiency of H₂O₂ photosynthesis.

Experimental

Materials

Melamine (99%), acetonitrile (99.9%), anthraquinone-2-carboxylic acid (AQ-COOH, 98%), silver nitrate (AgNO₃, 99.8%) and *p*-benzoquinone (BQ, 99%) were purchased from Sigma-Aldrich. Potassium chloride (KCl, 99.5%) was obtained from Alfa Aesar. Lithium chloride (LiCl, 99.8%), potassium bipthalate (PT, 99.9%), potassium iodide (KI, 99%) and β-carotene (96%) were received from Innochem. There was no further purification treatment on all of the chemicals.

Synthesis of CN-*x*% and BCN

1.0 g of melamine was ground with 10 g of KCl/LiCl mixed molten salts (with KCl mass fraction *x*%) for 30 min. The resulting mixture was then transferred to a muffle furnace and calcinated at 550 °C for 4 h at a heating rate of 5 °C min⁻¹. After cooling to ambient temperature, the crude product was repeatedly washed with hot deionized water, isolated by centrifugation, and dried at 60 °C overnight to obtain the CN-*x*% composite. For example, CN-70% was synthesized using 7.0 g of KCl, 3.0 g of LiCl and 1.0 g of melamine. For comparison, bulk g-C₃N₄ (BCN) was also prepared using the same procedure, but without the addition of molten salts.

Synthesis of AQ/CN-*x*%

A total of 0.4 g of CN-*x*% was dispersed in 40 mL of acetonitrile, followed by the addition of 0.1 g of anthraquinone-2-carboxylic acid (AQ-COOH). The mixture was sonicated for 30 min to ensure uniform dispersion and then stirred at 50 °C for 12 h. After cooling to room temperature, the product was collected by filtration, thoroughly washed with deionized water and ethanol to remove unreacted species, centrifuged, and vacuum-dried at 60 °C overnight to obtain the AQ/CN-*x*% composite. For comparison, AQ/BCN was synthesized *via* a similar method with BCN as the substrate.

Characterization

X-ray diffraction (XRD) patterns were acquired on a Bruker D8 Advance X diffractometer using Cu Kα radiation (λ = 1.5405 Å) at 40 kV. Scanning electron microscopy (SEM) was performed on a Hitachi S-4800 microscope. Transmission electron microscopy (TEM) and high-angle annular dark-field scanning TEM (HAADF-STEM) with energy-dispersive spectroscopy (EDS) mapping were conducted using Hitachi HT-7700 (100 kV) and JEOL JEM-ARM300F (300 kV) instruments, respectively. X-ray photoelectron spectroscopy (XPS) was carried out on an ESCALAB 250Xi spectrometer with Al-Kα excitation, and all binding energies were calibrated to the carbon C 1s peak at 284.8 eV. The ¹³C solid-state nuclear magnetic resonance (NMR) spectra were recorded on a Bruker Ascend TM 400WB. UV-Vis diffuse reflectance spectra were obtained using an Agilent Cary 7000 spectrophotometer equipped with an integrating sphere. Fourier transform infrared (FTIR) spectra were recorded using a Varian Excalibur spectrometer. Specific surface areas were



determined by Brunauer–Emmett–Teller (BET) analysis on a Quadrasorb SI-MP porosity analyzer. The steady-state photoluminescence spectra (PL) and time-resolved photoluminescence spectra (TRPL) were recorded at room temperature using time-correlated single-photon counting (Edinburgh Instruments, FLS1000). Femtosecond transient absorption spectroscopy (fs-TAS) measurements were performed employing a Coherent Vitora-Legend Elite-Helios system. For fs-TAS, aqueous catalyst dispersions (0.5 mg mL^{-1}) were sonicated for 30 min prior to measurement in a microcuvette at room temperature, using 320 nm excitation with probe wavelengths ranging from 320 to 800 nm. Electron paramagnetic resonance (EPR) measurement was conducted on an E500 spectrometer.

Photocatalytic activity evaluation

For photocatalytic testing, 20 mg of catalyst was dispersed in 20 mL of deionized water and sonicated for 30 min to ensure uniform dispersion. The photocatalytic reactions were conducted under visible light irradiation ($\lambda \geq 420 \text{ nm}$) using a 300 W Xenon lamp equipped with a 420 nm cut-off filter, in an open quartz tube under atmospheric conditions, with air serving as the O_2 source. Reaction aliquots (1 mL) were collected at 60 min intervals, centrifuged, and filtered through a $0.22 \mu\text{m}$ polyethersulfone (PES) membrane. The H_2O_2 concentration was determined by iodometry: 500 μL of 0.1 M PT solution and 500 μL of 0.4 M KI solution were sequentially added to 1 mL of the filtrate. After reaction in the dark for 1 h, the absorbance at 350 nm was measured using a UV-vis spectrophotometer. The H_2O_2 concentration was calculated from a pre-established calibration curve.

Computational details

To obtain a reasonable initial geometry for the catalyst model, we first conducted a conformational search using CREST 3.0.1 (ref. 19) with the GFN2- χTB^{20} tight-binding quantum chemical method. The lowest-energy conformer was then selected for further density functional theory (DFT) calculations. Geometry optimization was performed at the (u)B3LYP/def2SVP level,^{21–26} employing Becke's three-parameter hybrid functional combined with the Lee–Yang–Parr correlation functional. To accurately account for long-range dispersion interactions, Grimme's D3 (ref. 27) dispersion correction with Becke–Johnson (BJ)²⁸ damping was applied. Frequency calculations were conducted at the same level to confirm that a true minimum (*i.e.*, no imaginary frequencies) had been obtained on the potential energy surface. Time-dependent density functional theory (TD-DFT) calculations were performed at the TD-B3LYP-GD3(BJ)/def2SVP level of theory and the hole–electron analysis was based on these calculations using Multiwfn 3.8(dev).²⁹ All calculations were done using GAUSSIAN 16 C.01.

Results and discussion

Structure and composition

The synthesis route of AQ/CN- $x\%$ is illustrated in Fig. 1a. CN- $x\%$ (where x denotes the mass fraction of KCl in the molten salts)

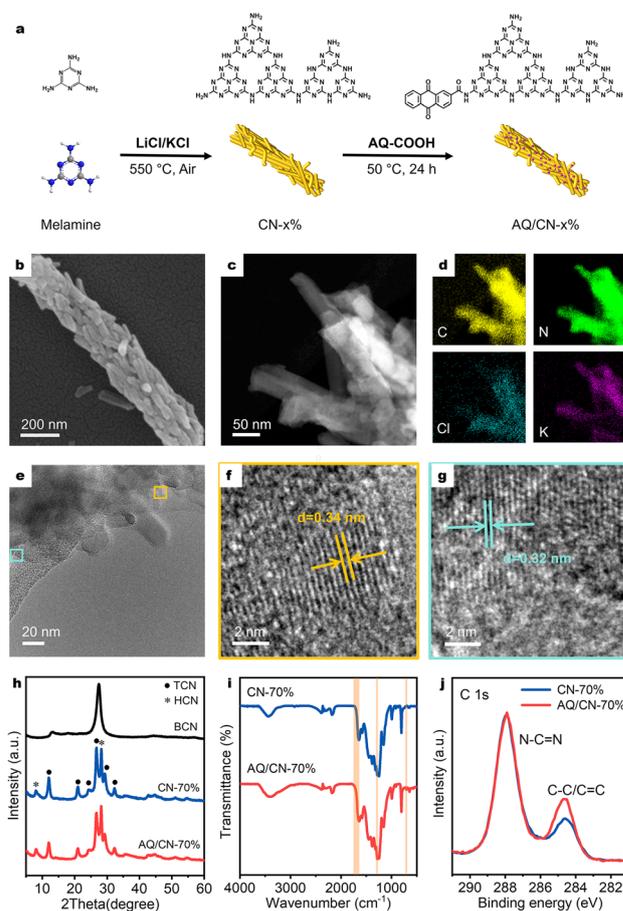


Fig. 1 (a) Schematic illustration of the synthesis route for AQ/CN- $x\%$, (b) SEM image of AQ/CN-70%, (c and d) HAADF-STEM and related EDX elemental mapping images of AQ/CN-70%, (e–g) TEM and HRTEM of AQ/CN-70%, with lattice fringes from regions marked by coloured boxes in (e), (h) XRD patterns of BCN, CN-70%, and AQ/CN-70%, (i) FTIR spectra of CN-70% and AQ/CN-70%, and (j) XPS spectra of C 1s for CN-70% and AQ/CN-70%.

containing both triazine (TCN) and heptazine (HCN) units was first prepared *via* molten salt-assisted thermal condensation. Subsequently, AQ units were introduced through an amidation reaction.¹⁸ As revealed by SEM and TEM images (Fig. 1b, S1a and b, and S2a–c), both CN-70% and AQ/CN-70% exhibit a hierarchical nanorod structure, with individual rods approximately 50 nm in diameter assembled into larger architectures. Compared to direct calcination, molten salt-assisted thermal condensation is more effective in promoting the reconstruction of carbon nitride crystals and the transformation of their aggregated states. HAADF-STEM and corresponding EDS mapping images (Fig. 1c and d) visually demonstrate the spatial distribution of C and N elements, along with trace amounts of K and Cl, within AQ/CN-70%. HRTEM images (Fig. 1e–g) further confirm the coexistence of TCN and HCN units in the hierarchical structure of AQ/CN-70%. Fourier-filtered analysis of selected regions resolved distinguishable lattice fringes with characteristic d -spacings of 0.34 and 0.32 nm, corresponding to the (002) crystallographic planes of TCN and HCN units, respectively.³⁰



XRD analysis further confirmed the crystalline phase composition of the catalysts. As depicted in Fig. 1h, BCN only showed characteristic peaks at 13.0° and 27.5° corresponding to the (100) and (002) planes of $g\text{-C}_3\text{N}_4$, respectively. In contrast, CN-70% displayed distinct diffraction peaks characteristic of both TCN and HCN units. Peaks at 12.0° , 20.9° , 24.2° , 29.1° , and 32.3° were assigned to the (100), (110), (200), (102), and (210) planes of the TCN structure, respectively, while the peak at 28.1° was attributed to the (002) plane of the HCN unit.^{30,31} Notably, AQ functionalization did not alter the crystalline phase composition of CN-70%. Moreover, the triazine/heptazine ratio could be tuned by adjusting the KCl content in the molten salts (Fig. S3). As KCl content decreased, the intensity of triazine-related peaks increased, indicating that KCl was more beneficial to inducing HCN unit formation, which was consistent with previous literature reports.³²

To confirm the successful AQ modification of CN-70%, FTIR and XPS were carried out. Compared with CN-70%, a new peak appeared at 698 cm^{-1} in AQ/CN-70% (Fig. 1i), corresponding to the bending vibration of C–H in the AQ group.^{33,34} Enhanced absorption peaks at 1278 cm^{-1} (C–N stretching vibration) and 1676 cm^{-1} (carbonyl C=O stretching vibration) further indicated the formation of an amide bond between AQ and CN-70%.³⁵ Additionally, Fig. S4 demonstrates that the TCN/HCN ratio can be precisely tuned by adjusting the KCl composition in the molten salts. In the ^{13}C solid-state NMR spectra (Fig. S5), the distinct signals observed at 163 ppm, 157 ppm, and 169 ppm for CN-70% and AQ/CN-70% are attributed to the $\text{CN}_2\text{-NH}_x$ carbon atoms in the carbon nitride skeleton, C– N_3 carbon atoms, and the $\text{-C}\equiv\text{N}$ carbon atoms, respectively.³⁶ For AQ/CN-70%, an additional set of new signals centered around 130 ppm emerged, which are ascribed to the aromatic C=C carbon atoms and carbonyl carbon atoms of the anthraquinone unit with slightly different chemical environments, confirming the successful incorporation of anthraquinone.³⁷ Furthermore, a new resonance appearing at approximately 180 ppm corresponds to the amide carbonyl carbon atom, indicating that the anthraquinone moiety is grafted onto the carbon nitride framework *via* an amide linkage.³⁸ This signal exhibits a notable downfield shift compared to the carbonyl carbon of a typical primary amide (170–175 ppm), which likely originates from the strong intermolecular hydrogen bonding interactions.³⁹

In the high-resolution XPS C 1s spectrum (Fig. 3b), characteristic peaks at 284.8, 286.1, and 288.1 eV correspond to C–C/C=C bonds, C– NH_x ($x = 1$ or 2) at the HCN ring edges, and sp^2 -hybridized carbon (N–C=N) within the conjugated system, respectively. As shown in Fig. 1j and Table S1, a notable increase in C–C/C=C content (from 22.7% in CN-70 to 27.8% in AQ/CN-70) was observed, which is attributed to the successful introduction of AQ units. The high-resolution N 1s spectra (Fig. S6) reveal peaks at 398.6, 400.2, 401.3, and 404.1 eV for all catalysts, assigned to sp^2 -hybridized nitrogen (C–N=C), tertiary nitrogen (N–(C)₃), N– H_x ($x = 1$ or 2), and $\pi\text{-}\pi^*$ satellite transitions, respectively. As shown in Fig. S6a and Table S2, a comparison between AQ/CN-70% and CN-70% reveals that the N– H_x group content in AQ/CN-70% significantly decreased, which is attributed to the consumption of edge amino groups during the

amidation reaction. Conversely, as shown in Fig. S6b, decreasing the KCl content led to a significant increase in the C– NH_x proportion, suggesting that TCN-rich structures, which possess a higher density of edge amino groups, are more favourable for AQ modification. In contrast, CN-80% and AQ/CN-80% exhibited minimal changes in C–C/C=C content (Fig. S6c), indicating that the limited availability of edge amino groups in CN-80% hindered effective AQ coupling.

Photocatalytic H_2O_2 production performance

The photocatalytic H_2O_2 production of the prepared photocatalysts in pure water was assessed under visible light irradiation ($\lambda \geq 420\text{ nm}$) in an open quartz tube under atmospheric conditions, with air serving as the O_2 source. As shown in Fig. 2a, pristine BCN exhibited a relatively low H_2O_2 generation rate of $12.0\ \mu\text{mol g}^{-1}\text{ h}^{-1}$, whereas CN-70%, featuring mixed TCN/HCN units, achieved a higher rate of $35.9\ \mu\text{mol g}^{-1}\text{ h}^{-1}$. This enhancement suggests that the TCN/HCN facilitates the separation of photogenerated charge carriers. Notably, after AQ modification, AQ/BCN and AQ/CN-70% demonstrated enhanced production rates of 25.6 and $165.3\ \mu\text{mol g}^{-1}\text{ h}^{-1}$, respectively. Meanwhile, the apparent quantum yield (AQY) of AQ/CN-70% was also measured and the results are presented in Table S3. Under 420 nm irradiation, the AQY reached 2.48%, suggesting that AQ/CN-70% exhibits high light-to-chemical energy conversion efficiency. Given that TCN units offer more modification sites, it is essential to further investigate the influence of TCN content on H_2O_2 production. As depicted in Fig. 2b, the H_2O_2 production rate gradually decreased as the KCl ratio decreased from 90% to 60%, which can be attributed to the superior photocatalytic H_2O_2 generation performance of HCN units compared to TCN units.⁴⁰ After AQ modification, all AQ/CN- $x\%$ samples exhibited enhanced catalytic activity to varying degrees, with AQ/CN-70% showing the highest H_2O_2 production rate (Fig. 2c). This can be explained by the fact that in CN-90%, the limited amino groups result in insufficient AQ modification. Conversely, when the TCN content is too high (CN-60%), the charge separation efficiency and O_2 reduction capability are suppressed. Thus, AQ/CN-70% with a balanced TCN/HCN ratio demonstrated optimal photocatalytic activity.

Photocatalytic performance tests under different atmospheres revealed the following order of H_2O_2 production: $\text{O}_2 > \text{air} > \text{Ar}$ (Fig. 2d), indicating that O_2 is an indispensable reactant for H_2O_2 synthesis. With prolonged illumination time, the cumulative H_2O_2 production continuously increased (Fig. 2e), demonstrating that the catalyst maintains good catalytic efficiency over extended periods. Furthermore, cycling stability tests showed only a slight decrease in photoactivity in two consecutive runs (Fig. 2f), which might be attributed to the protonation of key active sites leading to reduced catalytic performance.⁴¹ Notably, the photoactivity remained stable in subsequent cycles, confirming that the AQ units remained structurally intact throughout the catalytic process. To corroborate this observation, we examined the structural evolution of the material before and after photocatalytic reactions (Fig. S7). The XRD patterns (Fig. S7a) revealed no significant alteration in



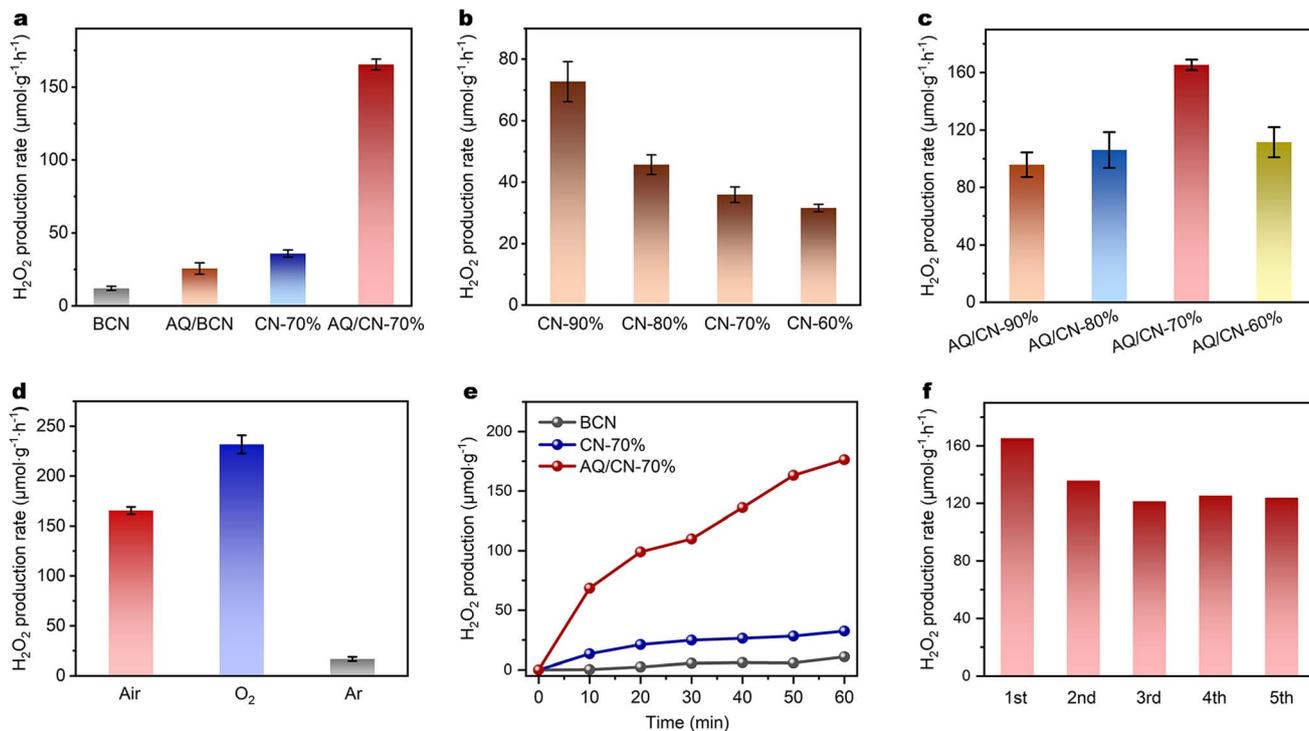


Fig. 2 (a) Photocatalytic H_2O_2 production rates of BCN, AQ/BCN, CN-70% and AQ/CN-70%, (b and c) H_2O_2 production rates of CN- $x\%$ and AQ/CN- $x\%$, (d) photocatalytic H_2O_2 production rate of AQ/CN-70% under varying atmospheres, (e) time-dependent H_2O_2 production kinetics of BCN, CN-70% and AQ/CN-70%, and (f) photocatalytic cycling stability of AQ/CN-70% for H_2O_2 production. Reaction conditions: 300 W Xenon light source equipped with a cut-off filter ($\lambda \geq 420$ nm); irradiation intensity: 300 mW cm^{-2} ; reaction temperature: 25°C .

the crystalline structure. Similarly, the FTIR spectra (Fig. S7b) confirmed the persistence of characteristic peaks associated with the AQ motifs. However, the intensities of the two peaks located at 1178 cm^{-1} and 998 cm^{-1} , which are assigned to the K^+ -coordinated $\text{C-N}=\text{C}$ bonds,^{42–44} were significantly reduced. This indicates that the performance degradation is caused by protonation. Electron microscopy revealed that the molten salt synthesis induced morphological changes, potentially influencing the specific surface area and consequently influencing photocatalytic performance. Nitrogen adsorption-desorption measurements (Fig. S8a–e) revealed that all samples exhibit similar pore size distributions. However, CN-70% and AQ/CN-70% exhibited comparable specific surface areas ($75.4 \text{ m}^2 \text{ g}^{-1}$ and $88.0 \text{ m}^2 \text{ g}^{-1}$, respectively), both of which are significantly larger than that of pristine BCN ($36.7 \text{ m}^2 \text{ g}^{-1}$). To disentangle surface area effects from intrinsic activity, we normalized the catalytic performance (Fig. S9). The results demonstrate that the substantial activity enhancement stems primarily from improved intrinsic activity due to AQ functionalization, rather than from merely an increased surface area.

Mechanism of photocatalytic hydrogen peroxide production

To elucidate the intrinsic mechanism underlying the enhanced photocatalytic performance, we comprehensively characterized the light absorption capability and band structure of the photocatalysts. UV-Vis diffuse reflectance spectroscopy (UV-Vis DRS) exhibited a slight red shift in the absorption edge and

enhanced absorption intensity in the $\lambda < 450 \text{ nm}$ region for CN-70% (Fig. 3a). After Kubelka–Munk transformation of the UV-Vis absorption spectra, the optical bandgap energies of BCN, CN-70% and AQ/CN-70% were determined to be 2.68 eV, 2.67 eV and 2.67 eV, respectively (Fig. 3b). Therefore, AQ modification did not alter the bandgap structure of CN-70%, indicating that the improvement in photocatalytic activity primarily originated from the AQ moiety. Fig. S10 presents the XPS valence band analysis results, showing that the valence band positions of BCN, CN-70%, and AQ/CN-70% relative to the normal hydrogen electrode (NHE) were 1.51 V, 1.64 V and 1.57 V, respectively. Based on the fundamental relationship $E_g = E_{\text{VB}} - E_{\text{CB}}$, the corresponding conduction band positions were calculated to be -1.17 V (BCN), -1.03 V (CN-70%) and -1.20 V (AQ/CN-70%), respectively. Fig. 3c illustrates the band structure diagram of the three photocatalysts, demonstrating all of them are suitable for driving oxygen reduction to generate H_2O_2 .⁴⁵

We then disclosed charge carrier separation and transfer dynamics through steady-state PL spectroscopy, TRPL spectroscopy and transient photocurrent measurements. PL spectra revealed a significant reduction in emission intensity for CN-70% compared to pristine BCN (Fig. 3d), indicating that the triazine rich $\text{g-C}_3\text{N}_4$ markedly enhanced charge separation efficiency. Notably, after AQ modification, the PL intensity of AQ/CN-70% further decreased, demonstrating that AQ functionalization could further promote the separation of photogenerated charges.⁴⁶ This finding was corroborated by TRPL



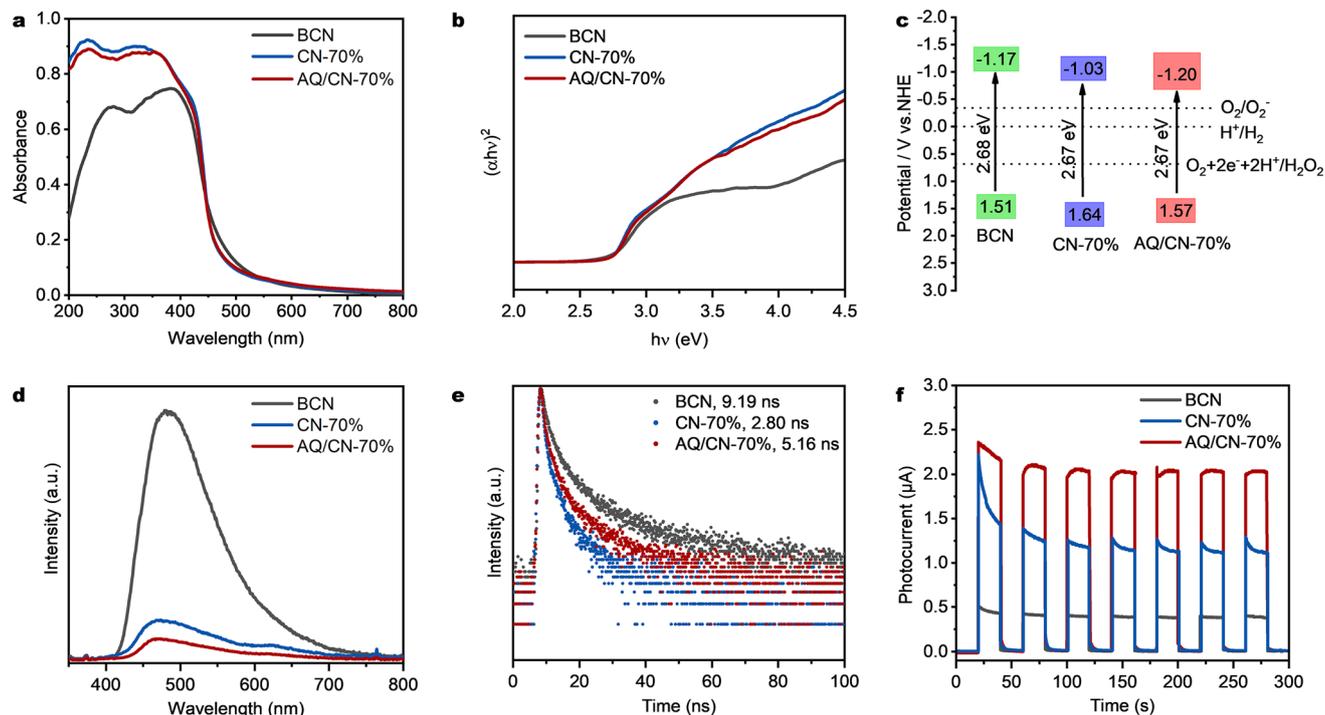


Fig. 3 (a) UV-Vis DRS and (b) Tauc plot of BCN, CN-70% and AQ/CN-70%, (c) band structure alignments of BCN, CN-70% and AQ/CN-70%, (d) steady-state PL spectra, (e) time-resolved transient PL spectra and (f) transient photocurrent responses of BCN, CN-70%, and AQ/CN-70%.

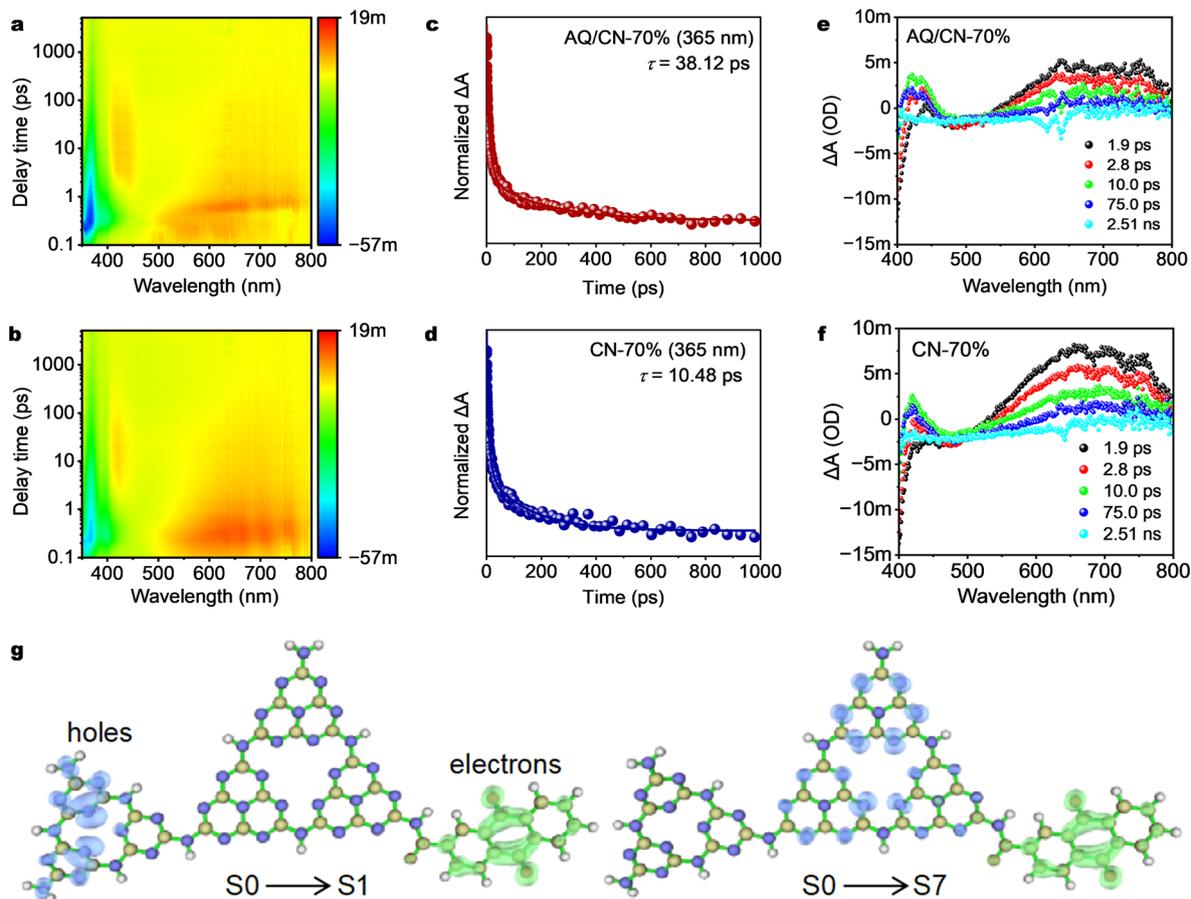


Fig. 4 2D mapping TA spectra of (a) AQ/CN-70% and (b) CN-70%, TA spectra signals on the fs–ns timescales of (c) AQ/CN-70% and (d) CN-70%, decay kinetic curves of (e) AQ/CN-70% and (f) CN-70%, and (g) the excited state information of AQ/CN-70% from DFT calculations.



measurements. As shown in Fig. 3e and S11, significantly shorter fluorescence lifetimes were observed for AQ/CN-70% and CN-70%, suggesting enhanced exciton dissociation into free charge carriers. Interestingly, AQ/CN-70% exhibited a slightly prolonged fluorescence lifetime compared to CN-70%, which can be attributed to the electron-withdrawing nature of the AQ units, which suppresses electron-hole recombination and extends charge carrier lifetimes.⁴⁷ Furthermore, transient photocurrent responses mirrored the photocatalytic activity trend (AQ/CN-70% > CN-70% > BCN). AQ/CN-70% generated the highest photocurrent density (Fig. 3f), directly evidencing that AQ modification effectively improved the separation efficiency of photogenerated charges.

To further investigate the dynamics of photoexcited states, we conducted fs-TAS analysis. As shown in Fig. 4a and b, the negative absorption peak at 350 nm corresponds to the ground-state bleaching (GSB) signal, originating from the pump-light

excitation of the photocatalyst. All samples exhibited strong GSB signals, which subsequently decayed rapidly due to electron-hole recombination.⁴⁸ The GSB signal intensity of AQ/CN-70% was significantly higher than that of CN-70%, indicating that AQ modification effectively increased the population of shallowly trapped electrons. Compared to CN-70%, AQ/CN-70% exhibited a notably prolonged GSB signal lifetime (Fig. 4c and d), demonstrating that the introduction of AQ units successfully suppressed electron-hole recombination.⁴⁵ The broad positive absorption feature observed in the 500–800 nm range represents the excited-state absorption (ESA) of shallowly trapped electrons.⁴⁹ The emergence of this positive signal reflects the enrichment of shallowly trapped electrons. Fig. S12a and c show that the lifetime of the excited-state electrons in AQ/CN-70% was significantly shortened, suggesting rapid transfer of excited-state electrons.⁵⁰ The characteristic transient absorption peaks in the 400–450 nm range originate from high-energy

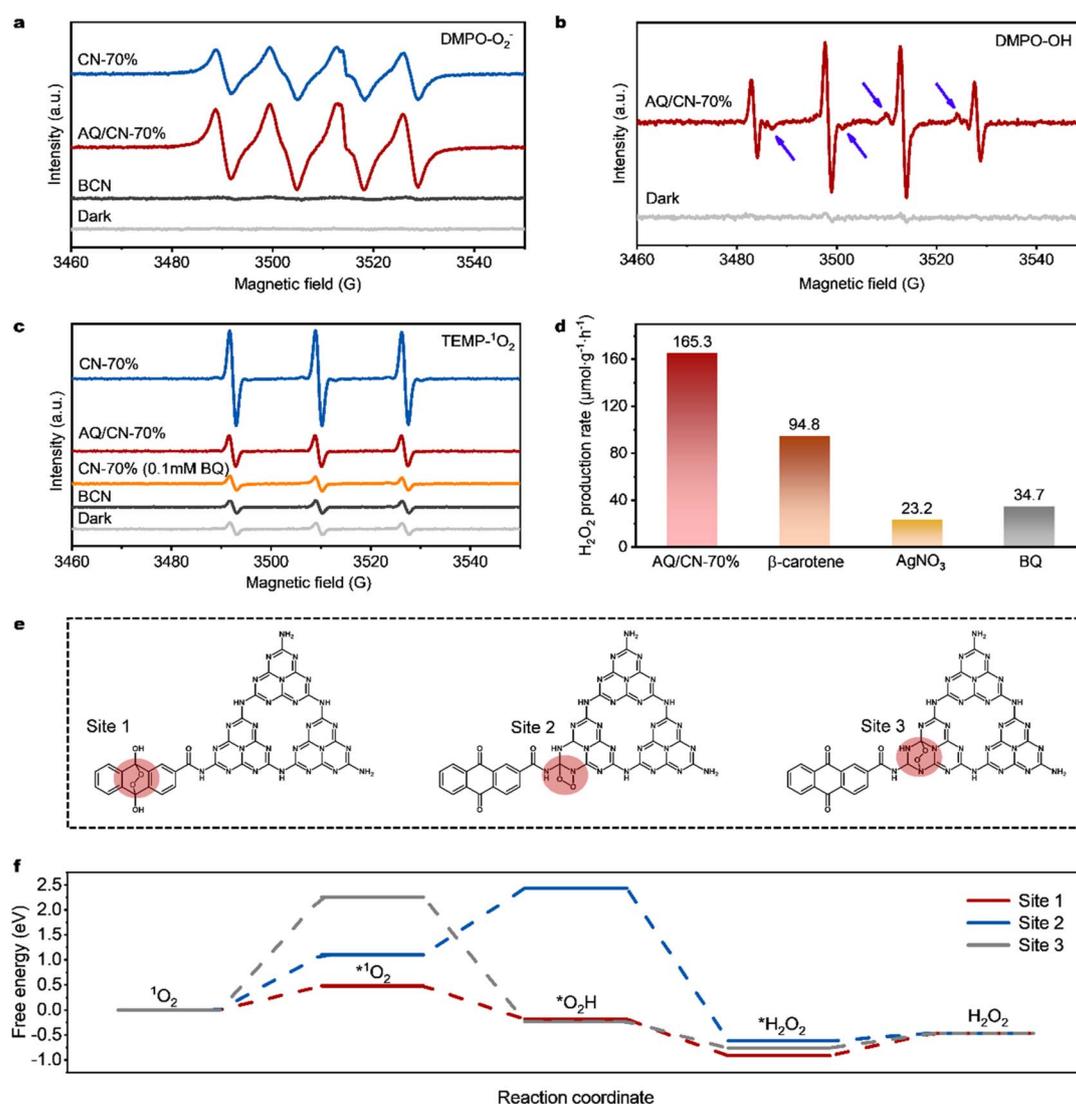


Fig. 5 (a) $\cdot\text{O}_2^-$, (b) $\cdot\text{OH}$ and (c) $^1\text{O}_2$ radical adduct signals in EPR spectra of BCN, CN-70%, AQ/CN-70% and CN-70% with the addition of BQ, (d) H_2O_2 production rates of AQ/CN-70% with different radical scavengers, (e) schematic illustration of potential $^1\text{O}_2$ reaction sites, and (f) free energy diagrams for H_2O_2 formation pathways at distinct reaction sites on AQ/CN-70%.



electronic transitions associated with Auger recombination of shallowly trapped charge carriers.^{51,52} Kinetic analysis (Fig. 4e and f and S13a and b) revealed that these narrow positive peaks appeared synchronously with the decay of the broad ESA signal (500–800 nm), exhibiting a characteristic lifetime of 300–400 ps (Fig. S12b and d), consistent with the emergence of high-energy Auger electrons. Notably, AQ/CN-70% not only retained the characteristic transitions observed in CN-70% at 400–450 nm but also exhibited a new peak at longer wavelengths. The appearance of this new peak suggests that the introduction of AQ units in AQ/CN-70% created a new energy level, which effectively captured high-energy electrons. This appearance of the new peak suggests that the introduction of AQ units in AQ/CN-70% created a new energy level, which effectively captured high-energy electrons generated *via* the Auger recombination process. DFT calculations (Fig. 4g and S14) further confirmed that the AQ group acts as an effective electron-withdrawing centre. Based on experimental and theoretical results, it was further demonstrated that the AQ units significantly enhanced charge separation due to their electron-withdrawing nature. Therefore, the AQ units not only promoted enhanced excited-state charge transfer but also suppressed electron–hole recombination, endowing AQ/CN-70% with exceptional photocatalytic performance.

By performing EPR during visible-light irradiation, we deciphered the radical-mediated reaction mechanism governing photocatalytic H₂O₂ production on the AQ/CN-70% catalyst. As shown in Fig. 5a, EPR measurements using 5,5-dimethyl-1-pyrroline *N*-oxide (DMPO) as a radical trapping agent in methanol solution detected no signal for AQ/CN-70% in the dark. Upon illumination, a characteristic $\cdot\text{O}_2^-$ signal appeared with significantly greater intensity than that of CN-70%, indicating that AQ modification enhanced charge separation and consequently improved the efficiency of O₂ reduction to $\cdot\text{O}_2^-$. Fig. 5b reveals the characteristic $\cdot\text{OH}$ signal peaks of AQ/CN-70% in deionized water using DMPO as a radical trapping agent during EPR testing, along with a quartet signal (blue arrow) attributed to $\cdot\text{OOH}$ radicals, suggesting a two-step single-electron oxygen reduction pathway for H₂O₂ production. Fig. 5c presents the EPR spectra of 2,2,6,6-tetramethyl-4-piperidone (TEMP)-trapped singlet oxygen (¹O₂). CN-70% exhibited the strongest ¹O₂ signal. When the superoxide radical ($\cdot\text{O}_2^-$) was scavenged by the quencher BQ, the ¹O₂ signal in CN-70% significantly weakened, confirming that the oxidation of $\cdot\text{O}_2^-$ by holes is the primary pathway for ¹O₂ generation. Therefore, the consumption of the intermediate $\cdot\text{O}_2^-$ for ¹O₂ formation competes with H₂O₂ production, adversely affecting photocatalytic performance. Fig. S15 reveals identical behaviour in AQ/CN-70%. Surprisingly, the ¹O₂ signal also sharply decreased after AQ modification (Fig. 5c). Given that the $\cdot\text{O}_2^-$ signal in AQ/CN-70% was enhanced under illumination, it is reasonable to conclude that the weakened ¹O₂ signal primarily results from the trapping and conversion of ¹O₂ by AQ units. To investigate the contributions of different reactive species to H₂O₂ generation, quenching experiments were conducted using specific scavengers: carotene (for ¹O₂), AgNO₃ (for e⁻) and BQ (for $\cdot\text{O}_2^-$). As shown in Fig. 5d, the addition of β -carotene reduced the H₂O₂ yield from 165.3 to

94.8 $\mu\text{mol g}^{-1} \text{h}^{-1}$, indicating that ¹O₂ contributes partially to H₂O₂ production. When AgNO₃ (e⁻ scavenger) and BQ ($\cdot\text{O}_2^-$ scavenger) were introduced, the H₂O₂ yield dropped drastically, confirming that electrons and $\cdot\text{O}_2^-$ remain the key reactive species. Based on the RDE measurements, the electron transfer number for AQ/CN-70% during the reaction was determined to be 1.89 (Fig. S16), indicating that AQ/CN-70% indeed produces hydrogen peroxide *via* a two-electron ORR pathway.

DFT calculations were further performed to evaluate the free energy of different ¹O₂ reaction sites on AQ/CN-70%. As shown in Fig. 5e, f and S17, the results revealed that, compared to the reaction sites in pristine carbon nitride (Site 2 and Site 3), the AQ site (Site 1) exhibited a significantly more favourable Gibbs free energy change (ΔG) for the formation of $\cdot^1\text{O}_2$ and $\cdot\text{OOH}$ intermediates. This indicates that the AQ site is more efficient in reducing ¹O₂ to H₂O₂. These findings further confirm that the AQ unit serves as the active site for converting ¹O₂ into H₂O₂. Therefore, AQ modification not only enhances charge separation but, more importantly, also transforms the generated ¹O₂ into the target product H₂O₂. This strategy ingeniously converts an otherwise competing side reaction into a favourable pathway for H₂O₂ production, fundamentally improving the efficiency of photocatalytic H₂O₂ synthesis.

Integrated experimental and theoretical evidence confirms the charge transfer dynamics in AQ/CN-70% (Fig. 6), characterized by hole accumulation at TCN units and electron localization on anthraquinone moieties. This spatial separation creates donor–acceptor (D–A) configurations, significantly enhancing charge separation efficiency and thereby boosting

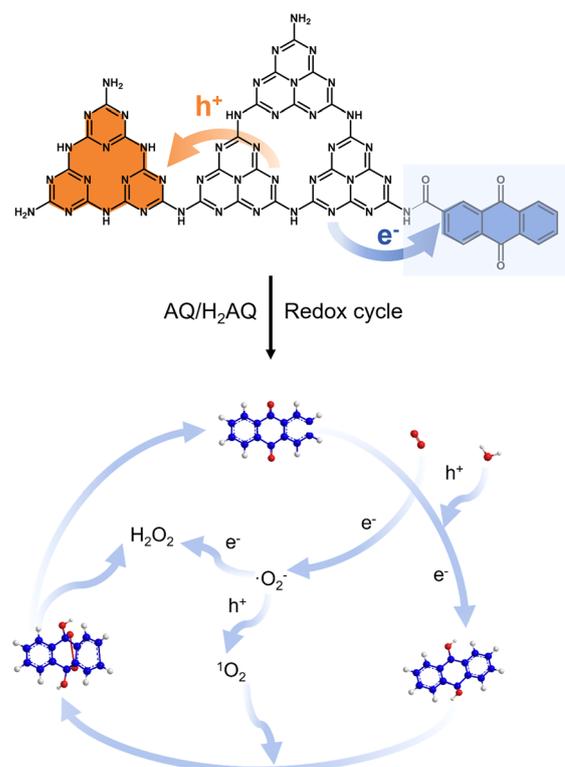


Fig. 6 Charge transfer and AQ/H₂AQ redox cycle mechanism for photocatalytic H₂O₂ synthesis on the AQ/CN-*x*% catalyst.



photocatalytic performance. EPR analysis reveals dual H_2O_2 production pathways: (1) conventional $\text{O}_2 \rightarrow \cdot\text{O}_2^- \rightarrow \text{H}_2\text{O}_2$ reduction and (2) a synergistic AQ/ H_2AQ cycle that converts detrimental $^1\text{O}_2$ into H_2O_2 through: (i) $\text{AQ} + 2\text{e}^- + 2\text{H}^+ \rightarrow \text{H}_2\text{AQ}$ (proton-coupled electron transfer) and (ii) $\text{H}_2\text{AQ} + ^1\text{O}_2 \rightarrow \text{H}_2\text{O}_2 + \text{AQ}$ (singlet oxygen scavenging). This dual-reduction mechanism circumvents kinetic limitations from competitive $\cdot\text{O}_2^-$ oxidation, substantially elevating H_2O_2 production yields.

Conclusions

In conclusion, an AQ/CN-70% photocatalyst was successfully prepared through molten salt-assisted calcination followed by amide condensation to anchor AQ units. The TCN rich $\text{g-C}_3\text{N}_4$ framework provided abundant modification sites. After AQ functionalization, the catalyst demonstrated significantly enhanced visible-light-driven H_2O_2 production in pure water, achieving a remarkable yield of $165.3 \mu\text{mol g}^{-1} \text{h}^{-1}$, which is 13.8 times higher than that of pristine $\text{g-C}_3\text{N}_4$. fs-TAS combined with theoretical calculations revealed that the AQ unit acts as an electron-trapping centre, substantially accelerating excited-state electron transfer while prolonging electron relaxation time. EPR spectroscopy and radical trapping experiments further confirmed that the introduction of AQ units effectively suppressed $^1\text{O}_2$ accumulation and instead facilitated its participation as an intermediate in H_2O_2 synthesis. This work presents a novel strategy for converting the competitive $^1\text{O}_2$ generation pathway into H_2O_2 production by AQ modification, offering valuable insights into simultaneously improving charge separation efficiency and enhancing H_2O_2 synthesis performance.

Author contributions

C. Zhou, G. Chen and X. Meng conceived the idea and designed the experiments. T. Zhang supervised the project. J. Li, L. Wang, H. Zhang and Q. Chen carried out the experiments and characterization. J. Li, C. Zhou, G. Chen and X. Meng wrote the manuscript. All the authors discussed the results and commented on the manuscript.

Conflicts of interest

There are no conflicts to declare.

Data availability

The data supporting this article are included in the SI. Supplementary information is available. See DOI: <https://doi.org/10.1039/d5ta05338b>.

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