

Cite this: *J. Mater. Chem. A*, 2025, 13, 10736

# Layered sodium vanadate ( $\text{NaV}_8\text{O}_{20}$ ) nanobelts: a new high-performing pseudocapacitive material for sodium-ion storage applications†

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Multifunctional layered nanostructures have attracted great attention for next-generation electrochemical supercapacitors and metal-ion batteries. Herein, we use a hydrothermal method to demonstrate the synthesis of 1D and layered sodium vanadate ( $\text{NaV}_8\text{O}_{20}$ ) nanobelts architecture. These  $\text{NaV}_8\text{O}_{20}$  nanobelts demonstrate outstanding electrochemical performance in supercapacitors (SCs) and sodium-ion batteries (SIBs). The possible formation mechanism of  $\text{NaV}_8\text{O}_{20}$  nanobelts is briefly discussed. Benefiting from the 1D and layered nanostructure, pre-inserted cations, significantly enhanced electrochemical conductivity, and high electroactive surface area, the prepared  $\text{NaV}_8\text{O}_{20}$  electrode material exhibited excellent charge storage capacity, favorable rate, and cyclic stability performance. The  $\text{NaV}_8\text{O}_{20}$  nanobelts displayed outstanding electrochemical characteristics, including  $676 \text{ F g}^{-1}$  of specific capacitance,  $45 \text{ W h kg}^{-1}$  of energy density and  $5224 \text{ W kg}^{-1}$  of power density. Additionally, on testing in Na-ion batteries, the  $\text{NaV}_8\text{O}_{20}$  nanobelts exhibit a discharge capacity of  $110 \text{ mA h g}^{-1}$  at  $10 \text{ mA g}^{-1}$  and retain  $\sim 52\%$  capacity after 100 cycles. Along with this, the galvanostatic intermittent titration technique (GITT) measurements reveal a high diffusion coefficient for  $\text{Na}^+$  ions, highlighting the efficient  $\text{Na}^+$  ions transportation within the  $\text{NaV}_8\text{O}_{20}$  structure. To our knowledge, this is the first report on the use of  $\text{NaV}_8\text{O}_{20}$  nanobelts for both SIBs and SCs, marking a significant contribution to the development of multifunctional materials for energy storage applications.

Received 4th December 2024  
Accepted 12th March 2025

DOI: 10.1039/d4ta08624d

rsc.li/materials-a

## 1. Introduction

The rapid advancement and proliferation of electrochemical energy storage (EES) technologies are reshaping modern life while

laying the foundation for a sustainable energy future. The growing demand for portable/wearable electronics, electric vehicles, and stationary energy storage systems has sparked significant global interest in EES technologies, particularly supercapacitors (SCs) and metal-ion batteries (MIBs). These devices have emerged as critical components for addressing the energy challenges of modern technology due to their distinct capabilities in energy storage and power delivery.<sup>1,2</sup> However, to address the growing requirements for energy storage systems, the performance of these EES devices must be significantly improved. The working electrode material is one of the most essential parameters for EES applications.<sup>3</sup> Therefore, enhancing the charge storage capacity of electrode materials has become a primary focus in recent EES research. One-dimensional (1D) and layered nanostructures have garnered significant scientific and technical attention for EES systems, attributed to their high electroactive surface area, structural stability, short charge carrier transport pathways, and favorable intercalation/de-intercalation processes.<sup>4</sup> Various 1D and layered nanostructures, including carbon nanostructures, conducting polymers, metal phosphides and metal chalcogenides, have been investigated as potential electrode materials (both cathode and anode) for SC and MIB applications.<sup>5–10</sup> Among these, transition metal oxides (TMOs) stand out as highly

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† Electronic supplementary information (ESI) available: S1: (a) chemicals and reagents, (b) formulae for  $C_{sp}$  calculation of working electrode and ASC device, and (c) characterization techniques, S2: XPS studies of  $\text{NaV}_8\text{O}_{20}$  nanobelts sample, S3: electrochemical measurements of carbon paper, bulk  $\text{V}_2\text{O}_5$  and  $\text{NaV}_8\text{O}_{20}$  electrode and Na ion battery, Table S1: crystallographic data of the  $\text{NaV}_8\text{O}_{20}$  powder obtained from Rietveld refinement, Table S2:  $C_{sp}$  values of  $\text{NaV}_8\text{O}_{20}$  electrode from CV curves, Table S3:  $C_{sp}$  values of  $\text{NaV}_8\text{O}_{20}$  electrode from GCD curves, Table S4:  $C_{sp}$ , energy and power density values of AC// $\text{NaV}_8\text{O}_{20}$  ASC coin cell device. See DOI: <https://doi.org/10.1039/d4ta08624d>

favorable electrodes for EES due to their multifarious oxidation states, high theoretical specific capacitance, diverse 1D/2D nanostructured morphologies, high electroactive surface area, and cost-effective fabrication processes.<sup>11–13</sup>

Transition metal vanadates, a class of bimetallic TMOs, have gathered great scientific attention as electrode materials for SCs and MIBs. Their appeal stems from their high surface area, low-dimensional and layered structures, excellent pseudocapacitive properties, enhanced electrical conductivity, and superior cyclic stability. These characteristics enable rapid and reversible surface redox reactions of electroactive species, making them highly suitable for advanced energy storage applications.<sup>14,15</sup> Currently, different binary transition metal vanadates such as  $\text{Co}_3\text{V}_2\text{O}_8$ ,  $\text{FeVO}_4$ ,  $\text{MnV}_2\text{O}_8$ ,  $\text{Ni}_3\text{V}_2\text{O}_8$ , and  $\text{Cu}_3\text{V}_2\text{O}_8$  have proven their ability in SCs and MIBs applications owing to their wide working potential window, high specific capacitance, and energy storage capacity compared to their pristine metal oxide counterparts.<sup>16–20</sup> For example, quasi-cuboidal  $\text{CoV}_2\text{O}_6$  has been shown to exhibit a specific capacitance of  $223 \text{ F g}^{-1}$  at  $1 \text{ A g}^{-1}$ , along with remarkable cyclic stability of 123% after 15 000 cycles, making it suitable for SC applications.<sup>21</sup> Similarly, 3D flower-like  $\text{Ni}_3\text{V}_2\text{O}_8$ @CNTs have proven effective in enhancing energy storage capacity due to their high specific capacitance of  $1054 \text{ F g}^{-1}$  at  $1 \text{ A g}^{-1}$  and retain high cyclic stability of 89% after 10 000 cycles.<sup>22</sup> Pseudocapacitive layered  $\text{Fe}_5\text{V}_{15}\text{O}_{39}(\text{OH})_9 \cdot 9\text{H}_2\text{O}$  nanosheets exhibit significant potential as a cathode material for lithium-ion batteries, achieving a high reversible capacity of  $350 \text{ mA h g}^{-1}$  at  $100 \text{ mA g}^{-1}$ , indicating its capability for long-term cycling stability.<sup>23</sup> Similarly,  $\text{Zr-NH}_4\text{V}_4\text{O}_{10}$  has been identified as a promising cathode material for sodium-ion batteries, achieving a high specific capacity of  $342 \text{ mA h g}^{-1}$  at  $100 \text{ mA g}^{-1}$ , positioning it as an appealing option for large-scale energy storage systems.<sup>24</sup> However, despite these advancements, several challenges persist, including the complex synthesis procedures, use of toxic and hazardous materials and high cost of precursor materials, which hinder the feasibility of these materials for widespread industrial applications.<sup>15</sup> Due to its numerous significant benefits, the hydrothermal method is the most frequently utilized for synthesizing metal vanadates. This approach enables accurate regulation of the morphology and particle size of resultant nanostructures, which is essential for enhancing the electrochemical performance of energy storage devices. This process is environmentally friendly and scalable, capable of being conducted at relatively low temperatures and pressures, contributing to its energy efficiency and cost-effectiveness. Furthermore, the hydrothermal approach facilitates the creation of well-crystallised nanostructures that exhibit high purity, thereby improving their applicability in energy storage applications. Considering these aspects, low-cost and environment-friendly alkali metals (Li, Na, K) can be appropriate alternatives to toxic and expensive metals for developing metal vanadate materials. However, alkali metal-inserted vanadate materials have been less explored for electrochemical SC and MIB applications.

In this study, we address these above-mentioned challenges by developing a simple, cost-effective hydrothermal route to prepare the 1D and layered structural sodium vanadate ( $\text{NaV}_8\text{O}_{20}$ ) nanobelts without any post-annealing treatment. Sodium vanadate is

particularly gaining attention due to its pre-inserted sodium ions, which enhances electrochemical performance and its suitability as a low-cost charge storage material for SCs and sodium-ion battery (SIB) applications. To the best of our knowledge, no prior publication has been reported that explored the electrochemical SCs and sodium-ion battery (SIB) applications of the 1D and layered  $\text{NaV}_8\text{O}_{20}$  structures. Owing to the unique 1D layered structure and nanobelts morphology features, the  $\text{NaV}_8\text{O}_{20}$  electrode displayed outstanding electrochemical features with high specific capacitance ( $C_{\text{sp}}$ ) and cyclic stability. The  $\text{NaV}_8\text{O}_{20}$  material was investigated using various electrochemical and physicochemical characterization techniques to evaluate its structural, morphological and electrochemical characteristics. We have also successfully demonstrated the practical applicability of  $\text{NaV}_8\text{O}_{20}$  nanobelts in electrochemical applications, particularly in solid-state asymmetric SCs, highlighting their potential for advanced EES systems.

## 2. Experimental section

### 2.1 Synthesis of $\text{NaV}_8\text{O}_{20}$ nanobelts

The details of chemicals and reagents are provided in ESI, S1(a).<sup>†</sup> The chemicals and reagents used for all the experiments were of analytical grades. A hydrothermal method was used to synthesize  $\text{NaV}_8\text{O}_{20}$  material (Fig. 1a). First, 0.75 g of  $\text{V}_2\text{O}_5$  powder was dispersed in 40 ml of double distilled water and ultrasonicated for 90 min. Then, with continuous stirring, 1 M  $\text{Na}_2\text{SO}_4$  (30 ml) aqueous solution was gradually added to the  $\text{V}_2\text{O}_5$  suspension. After 30 min, the mixture was transferred into a 100 ml Teflon-lined stainless-steel autoclave and heated to  $180^\circ\text{C}$  for 48 h. After hydrothermal treatment, the obtained greenish-colored product was washed several times with double distilled water. Finally, the obtained product was dried at  $150^\circ\text{C}$  in a vacuum oven.

### 2.2 Preparation of working electrode, AC// $\text{NaV}_8\text{O}_{20}$ asymmetric SC and Na-ion battery

The schematic representation of the working electrode fabrication *via* the slurry casting method is illustrated in Fig. 1b. Initially, a homogeneous slurry was prepared by mixing  $\text{NaV}_8\text{O}_{20}$  nanobelts (active material), polyvinylidene difluoride (PVDF) as a binder, and activated carbon (AC) as a conductive additive in a weight ratio of 8:1:1, using *N*-methyl-pyrrolidone (NMP) as the solvent. This slurry was then uniformly coated onto a Toray carbon paper substrate. The coated electrode was subsequently dried under vacuum at  $150^\circ\text{C}$  for 24 hours. The final electrode had an active surface area of  $1 \times 1 \text{ cm}^2$ , with a mass loading of  $\sim 1.2 \text{ mg}$ .

Fig. 1c and d depicts the schematic for the device manufacture processes for SC and SIB coin cells of the  $\text{NaV}_8\text{O}_{20}$  nanobelts material. As shown in Fig. 1c, the AC// $\text{NaV}_8\text{O}_{20}$  asymmetric SC (ASC) device was fabricated using  $\text{NaV}_8\text{O}_{20}$  nanobelts as cathode ( $\sim 1.2 \text{ mg}$  of mass loading) and AC as an anode ( $\sim 2 \text{ mg}$  of mass loading) materials prepared onto a circular 16 mm carbon paper current collector. These electrodes were separated using a cellulose membrane. The PVA- $\text{NaClO}_4$  gel electrolyte was prepared by mixing 2 g  $\text{NaClO}_4$  and 2 g PVA in 25 ml DDW and then heated at  $80^\circ\text{C}$  for 3 h. Finally, the AC// $\text{NaV}_8\text{O}_{20}$  asymmetric SC device was



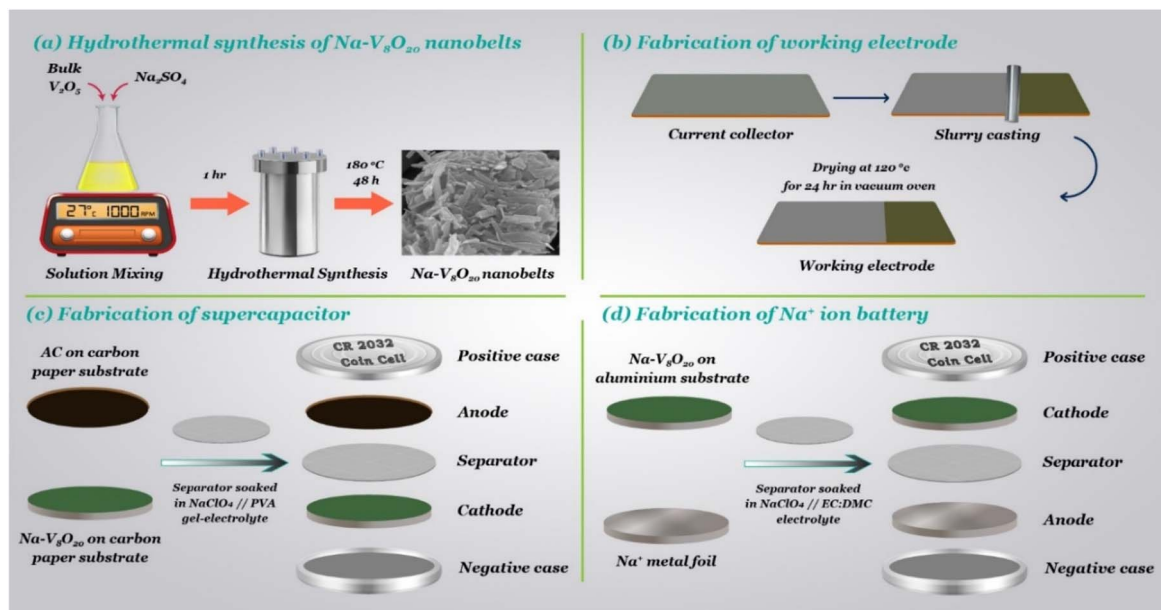


Fig. 1 Schematic representation of (a) hydrothermal synthesis of NaV<sub>8</sub>O<sub>20</sub> nanobelts, (b) preparation of working electrode, (c) fabrication of ASC device and (d) fabrication of Na<sup>+</sup> ion battery.

fabricated with a CR2032 coin cell using the PVA–NaClO<sub>4</sub> gel electrolyte.

The sodium-ion storage performance of NaV<sub>8</sub>O<sub>20</sub> nanobelts as the active cathode material was investigated. The electrode was fabricated by combining NaV<sub>8</sub>O<sub>20</sub>, PVDF, and AC with a weight ratio of 8 : 1 : 1 and thoroughly mixed in NMP solvent. The prepared homogeneous slurry was uniformly coated onto a current collector (aluminium foil) using the doctor blade method and subsequently dried under a vacuum oven at 100 °C for 12 hours. The dried film was then cut into 16 mm diameter discs with a mass loading of approximately 0.745 mg cm<sup>−2</sup>. Sodium metal was utilized as the counter and reference electrode, while a 19 mm diameter glass-fiber Whatman separator was employed. The 2032-type coin cells were assembled in an argon-filled glove box, where oxygen and moisture levels were controlled at 0.1 ppm. A 1 M solution of NaClO<sub>4</sub> in ethylene carbonate (EC) and diethyl carbonate (DEC) (1 : 1 v/v), with the addition of 5% fluoroethylene carbonate (FEC), was used as the electrolyte. A schematic representation of the Na<sup>+</sup> ion battery assembly process is provided in Fig. 1d.

### 2.3 Electrochemical measurements

All the electrochemical measurements of electrodes and fabricated devices, such as cyclic voltammetry (CV), galvanostatic charging–discharging (GCD), electrochemical impedance spectroscopy (EIS), rate capability, and cycling stability tests were conducted on the IVIUM-Vertex 1A electrochemical workstation (Ivium Technologies, Netherlands).

### 2.4 Physicochemical characterization of 1D and layered nanostructured NaV<sub>8</sub>O<sub>20</sub> nanobelts

The physicochemical characteristics of the synthesized material were investigated using various techniques, including X-ray

diffraction (XRD), Brunauer–Emmett–Teller (BET) surface area analysis, field emission scanning electron microscopy (FESEM), transmission electron microscopy (TEM), and X-ray photoelectron spectroscopy (XPS). Further details regarding these methods are provided in the ESI, S1(b).†

## 3. Results and discussion

### 3.1 XRD studies

XRD characterization was carried out to identify the crystal structure and phase purity of the NaV<sub>8</sub>O<sub>20</sub> material. Fig. 2a shows the comparative XRD pattern of bulk V<sub>2</sub>O<sub>5</sub> and NaV<sub>8</sub>O<sub>20</sub> powders. The diffraction pattern of bulk V<sub>2</sub>O<sub>5</sub> powder displays sharp and intense diffraction peaks, which match well to α-V<sub>2</sub>O<sub>5</sub> phase (JCPDS no. 77-2418) with an orthorhombic crystal structure. By comparing both the diffraction patterns, it is clear that, after hydrothermal treatment, the V<sub>2</sub>O<sub>5</sub> phase is converted into the newly formed NaV<sub>8</sub>O<sub>20</sub> phase. In the XRD pattern of NaV<sub>8</sub>O<sub>20</sub> powder, all the observed peaks are consistent with the characteristic peaks of the NaV<sub>8</sub>O<sub>20</sub> phase (JCPDS no. 45-1363) with a monoclinic crystal structure. Furthermore, no additional diffraction peaks related to V<sub>2</sub>O<sub>5</sub> or any other sodium vanadate phases were observed, implying the high phase purity of the NaV<sub>8</sub>O<sub>20</sub> phase. However, the NaV<sub>8</sub>O<sub>20</sub> phase displays relatively low-intensity and broad diffraction peaks compared with bulk V<sub>2</sub>O<sub>5</sub>, corresponding to the small crystallite size and low crystallinity of the NaV<sub>8</sub>O<sub>20</sub> nanobelts material. The intense diffraction from the (001) plane illustrates the preferential orientation of the NaV<sub>8</sub>O<sub>20</sub> phase in the *c* direction, which possesses the layered crystal structure arrangement.<sup>25</sup> The interlayer spacing (*d*-value) estimated from the (001) diffraction peak was ~10 Å. This large interlayer spacing is favorable for accommodating Na<sup>+</sup> ions during electrochemical processes.





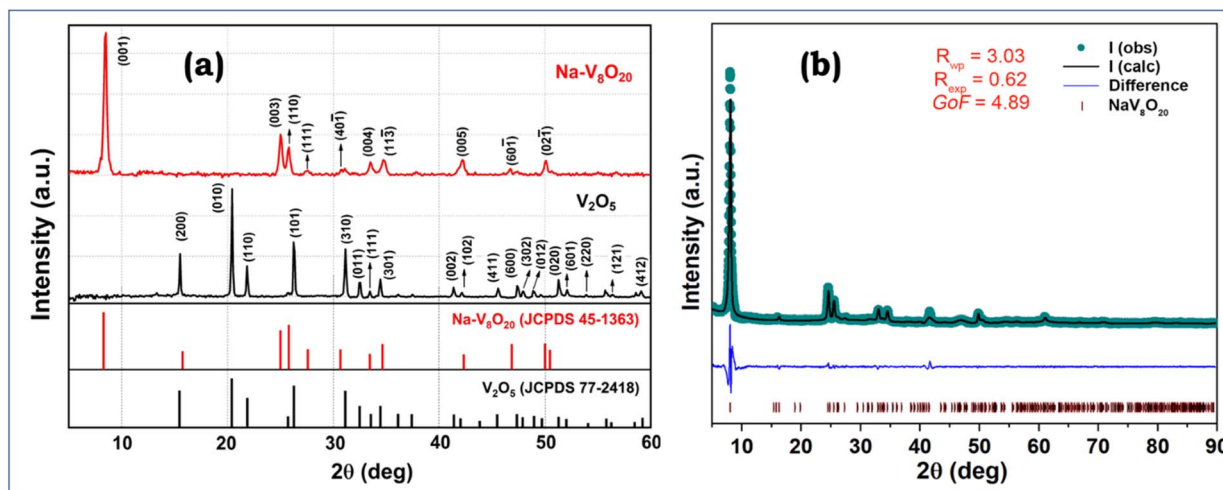


Fig. 2 (a) XRD patterns of  $V_2O_5$  and  $NaV_8O_{20}$  powder samples and (b) XRD refinement analysis of  $NaV_8O_{20}$  powder sample.

Rietveld refinement of the  $NaV_8O_{20}$  phase was conducted to identify the position of the sodium ions in the  $NaV_8O_{20}$  phase. Fig. 2b depicts the Rietveld refinement analysis of the  $NaV_8O_{20}$

nanobelts, performed using the MAUD software. The XRD pattern of the  $NaV_8O_{20}$  sample was fitted using the standard crystallography information file (CIF) consisting of

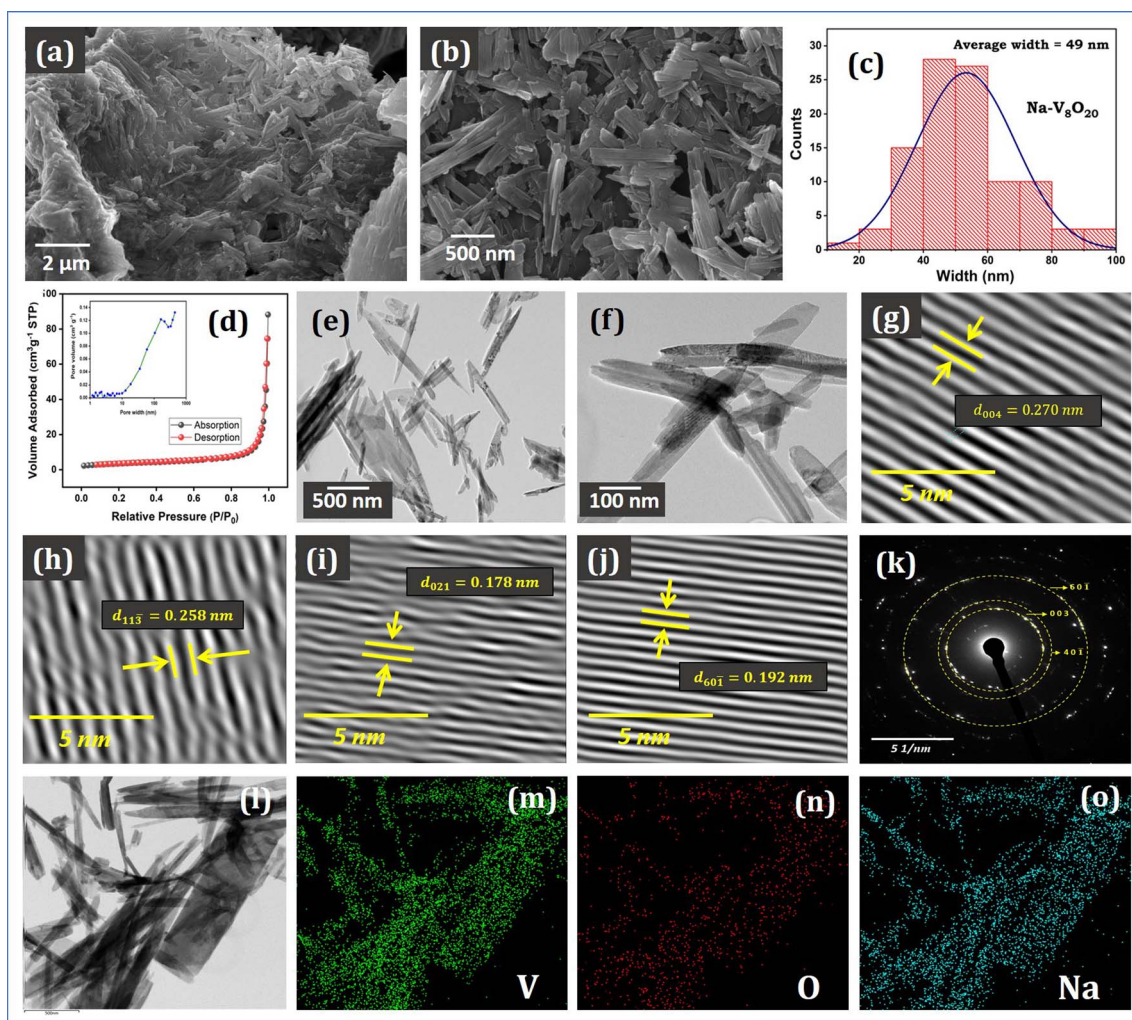


Fig. 3 (a and b) FESEM images, (c) histogram for average width; (d)  $N_2$  adsorption/desorption isotherms; (e and f) TEM images, (g–j) HR-TEM images, (k) SAED pattern and (l–o) elemental mapping of  $NaV_8O_{20}$  nanobelts.

(Na,Ca)(V,Fe)<sub>8</sub>O<sub>20</sub>·*n*H<sub>2</sub>O and JCPDS no. 45-1363.<sup>26</sup> By replacing the Ca and Fe atomic sites with corresponding amounts of Na and V, the refined analysis resulted in good agreement factors ( $R_{wp} = 3.03$ ,  $R_{exp} = 0.62$  and goodness of fit,  $GoF = 4.89$ ). The parameters that are refined are (1) microstructure, (2) basic phase parameters, (3) background and scale parameters, and (4) crystal structure (lattice parameter, occupancy, fractional coordinate, and isotropic displacement factor,  $B_{iso}$ ). The refined lattice parameters are:  $a = 11.74(2)$  Å,  $b = 3.6600(6)$  Å, and  $c = 11.117(5)$  Å with monoclinic phase,  $C2/m$  space group and  $\beta$  angle value =  $102.4(1)^\circ$ . The refinement data are mentioned in the ESI, Table S1.†

### 3.2 FE-SEM, BET and TEM studies

The morphological properties of NaV<sub>8</sub>O<sub>20</sub> material were elucidated by FESEM techniques (Fig. 3a and b). The low-resolution FESEM image (Fig. 3a) shows a network of uniformly distributed nanobelts adjoining each other to establish the self-assembled architecture of NaV<sub>8</sub>O<sub>20</sub> nanobelts. This 1D and densely attached nanobelt morphology will be favorable for electron transportation and will help to enhancing the diffusion of Na<sup>+</sup> ions by serving maximum electroactive sites during the electrochemical interactions.<sup>27,28</sup>

Fig. 3b represents the high-resolution FESEM image of the NaV<sub>8</sub>O<sub>20</sub> nanobelts. The width of NaV<sub>8</sub>O<sub>20</sub> nanobelts was measured using ImageJ software and the Gaussian function was used to calculate the average width of nanobelts. The calculated average width (Fig. 3c) of the NaV<sub>8</sub>O<sub>20</sub> nanobelt was ~49 nm. Additionally, to identify the surface area and pore size of the NaV<sub>8</sub>O<sub>20</sub> nanobelts sample, N<sub>2</sub> adsorption/desorption isotherm analysis was used. Fig. 3d shows the mesopores nature of NaV<sub>8</sub>O<sub>20</sub> nanobelts with a specific surface area of 12.41 m<sup>2</sup> g<sup>-1</sup> and pore size of ~43.3 nm (inset of Fig. 3d). Furthermore, crystallinity, morphological and elemental analysis of the NaV<sub>8</sub>O<sub>20</sub> nanobelts sample was investigated through TEM analysis. The low and high-magnification TEM images (Fig. 3e and f) display the nanobelts-like morphology of the NaV<sub>8</sub>O<sub>20</sub> material. The corresponding HRTEM images (Fig. 3g–i) depict the well-observable lattice fringes. The measured *d*-spacings of 0.22, 0.45, and 0.36 nm correspond to the (204), (201), and (003) lattice planes of the NaV<sub>8</sub>O<sub>20</sub> nanobelts. The selected area electron diffraction (SAED) pattern (Fig. 3k) further confirms the crystallinity of the NaV<sub>8</sub>O<sub>20</sub> nanobelts, revealing their polycrystalline nature. Furthermore, the elemental mapping images

(Fig. 3l–o) show a uniform spreading of Na, V, and O elements all over the NaV<sub>8</sub>O<sub>20</sub> nanobelts surface.

The X-ray photoelectron spectroscopy analysis of the NaV<sub>8</sub>O<sub>20</sub> nanobelts sample is provided in the ESI, S2.†

### 3.3 Formation mechanism of NaV<sub>8</sub>O<sub>20</sub> nanobelts

The possible formation mechanism of the NaV<sub>8</sub>O<sub>20</sub> nanobelts (Fig. 4) can be proposed based on their 1D nanobelt morphology and layered crystal structure, as confirmed by structural and morphological characterizations.

A hydrothermal synthesis route was used to prepare 1D and layered NaV<sub>8</sub>O<sub>20</sub> nanobelts. In the typical synthesis procedure, bulk V<sub>2</sub>O<sub>5</sub> powder (vanadium precursor) was pre-treated with ultrasonication in double distilled water to break bulk vanadium particles into smaller particles. This ultrasonication process promotes the dissociation and formation of free vanadium ionic species such as decavanadate ([V<sub>10</sub>O<sub>28</sub>]<sup>6-</sup>), polyoxovanadate (V<sub>8</sub>O<sub>20</sub><sup>4-</sup>), vanadyl ions (VO<sup>2+</sup>), hydrated vanadium species *etc.*, because of partial reduction and condensation of vanadium.<sup>29,30</sup>

Here, the presence of water plays a vital role in the hydrolysis of V<sub>2</sub>O<sub>5</sub>. When Na<sub>2</sub>SO<sub>4</sub> is added to the ionic solution of vanadium, the Na<sup>+</sup> ions interact with V<sub>8</sub>O<sub>20</sub><sup>4-</sup> ions, forming hydrated sodium vanadate complexes. Here, the free sulfate ions (SO<sub>4</sub><sup>2-</sup>) do not participate directly in the formation of the final product and water molecules (crystal water) help to stabilize these complexes by maintaining charge balance, leading to the formation of the NaV<sub>8</sub>O<sub>20</sub>·2H<sub>2</sub>O (complex), where water integrates sodium ions into the polyvanadate structure.

During the hydrothermal treatment, the system is subjected to high temperature and pressure, promoting dehydration and crystallization. The removal of crystal water during this step drives the organization of sodium vanadate species into well-defined and layered nanobelt structures. The dehydration process facilitates the reorganization of the vanadate layers, promoting anisotropic and preferential growth NaV<sub>8</sub>O<sub>20</sub> nanobelts along the *c* direction, confirmed by XRD with an intense peak at 8.8° with Miller indices (001).

Afterward, the FESEM image reveals that the NaV<sub>8</sub>O<sub>20</sub> nanobelts adopt a self-assembled morphology. This may be due to the removal of crystal water during crystallization, facilitating the formation of smooth, well-defined nanobelts. The self-assembly process is further driven by the depreciation of structural energy during crystallization, with sodium ions playing a role in charge stabilization.<sup>31</sup>

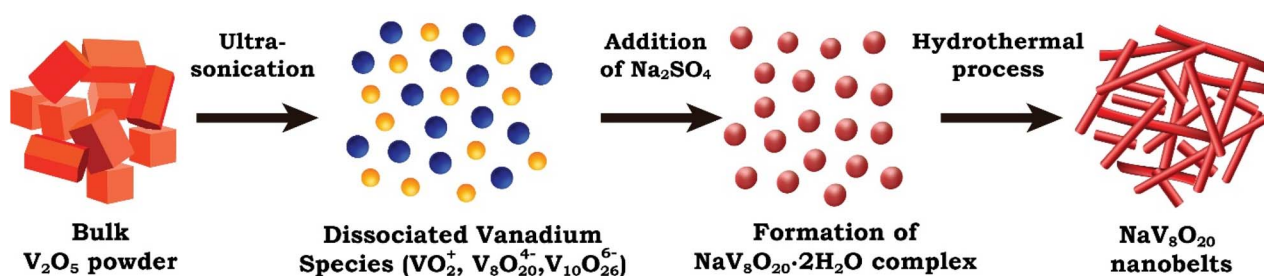


Fig. 4 Schematic representation for the possible formation mechanism of NaV<sub>8</sub>O<sub>20</sub> nanobelts.





## 4. Electrochemical investigation

### 4.1 Electrochemical performance of NaV<sub>8</sub>O<sub>20</sub> nanobelts electrode

The electrochemical charge storage features of the NaV<sub>8</sub>O<sub>20</sub> electrode material for supercapacitor studies were evaluated in 1 M Na<sub>2</sub>SO<sub>4</sub> electrolyte within a 0 to 0.6 V voltage window. The CV curves of the NaV<sub>8</sub>O<sub>20</sub> electrode were recorded for various scan rates ranging from 2.5 to 50 mV s<sup>-1</sup>, as shown in Fig. 5a. In Fig. 5a, the CV curves of the NaV<sub>8</sub>O<sub>20</sub> electrode exhibit a typical extrinsic pseudocapacitive nature consisting of a nearly rectangular shape with broad oxidation and reduction peaks. This extrinsic pseudocapacitive nature mainly arises from surface faradaic charge transfer reactions due to the intercalation or diffusion of charge-compensating electrolyte ions.<sup>32</sup> However, this nature is well retained despite decreasing the scan rate up to 2.5 mV s<sup>-1</sup>, implying outstanding rate performance of the NaV<sub>8</sub>O<sub>20</sub> electrode in Na<sub>2</sub>SO<sub>4</sub> aqueous solutions. From the CV curves, the  $C_{sp}$  values for different scan rates of the NaV<sub>8</sub>O<sub>20</sub> electrode were calculated using eqn (E1) (ESI, S1(b)†) and are given in the ESI, Table S2.†

The CV curves of the NaV<sub>8</sub>O<sub>20</sub> electrode reveal a hybrid charge storage behaviour, combining pseudocapacitive and

electric double-layer (EDLC) characteristics. The Trasatti method is employed to understand the charge storage mechanism of the NaV<sub>8</sub>O<sub>20</sub> nanobelts. This method helps distinguish between the contributions of surface-capacitive and diffusion-controlled processes. According to Trasatti, the overall capacitance arises from a blend of these two mechanisms: surface capacitive reactions and diffusion-driven charge storage.<sup>33,34</sup> Fig. 5b depicts the Trasatti plot, plotted using  $C_{sp}$  against the square root of the scan rate. The y-axis intercept of the linear fit represents the surface-controlled contribution. Based on this intercept, the proportion of surface capacitive and diffusion-controlled charge storage components was calculated across various scan rates. These results are displayed in Fig. 5c as a histogram. This histogram shows that the surface capacitive (EDLC) contribution has the dominant charge storage mechanism over the diffusion-controlled contribution. By increasing the scan rate from 2.5 to 50 mV s<sup>-1</sup>, the surface capacitive contributions improved from 56 to 85% with no noticeable CV curve distortion detected, revealing a better rate performance of the NaV<sub>8</sub>O<sub>20</sub> electrode. Furthermore, at lower scan rates (2.5 to 20 mV s<sup>-1</sup>), broad redox peaks and quasi-rectangular CV curves were observed (ESI, Fig. S3(d)†). This is due to surface faradaic reactions, clarifying the extrinsic pseudocapacitive nature.

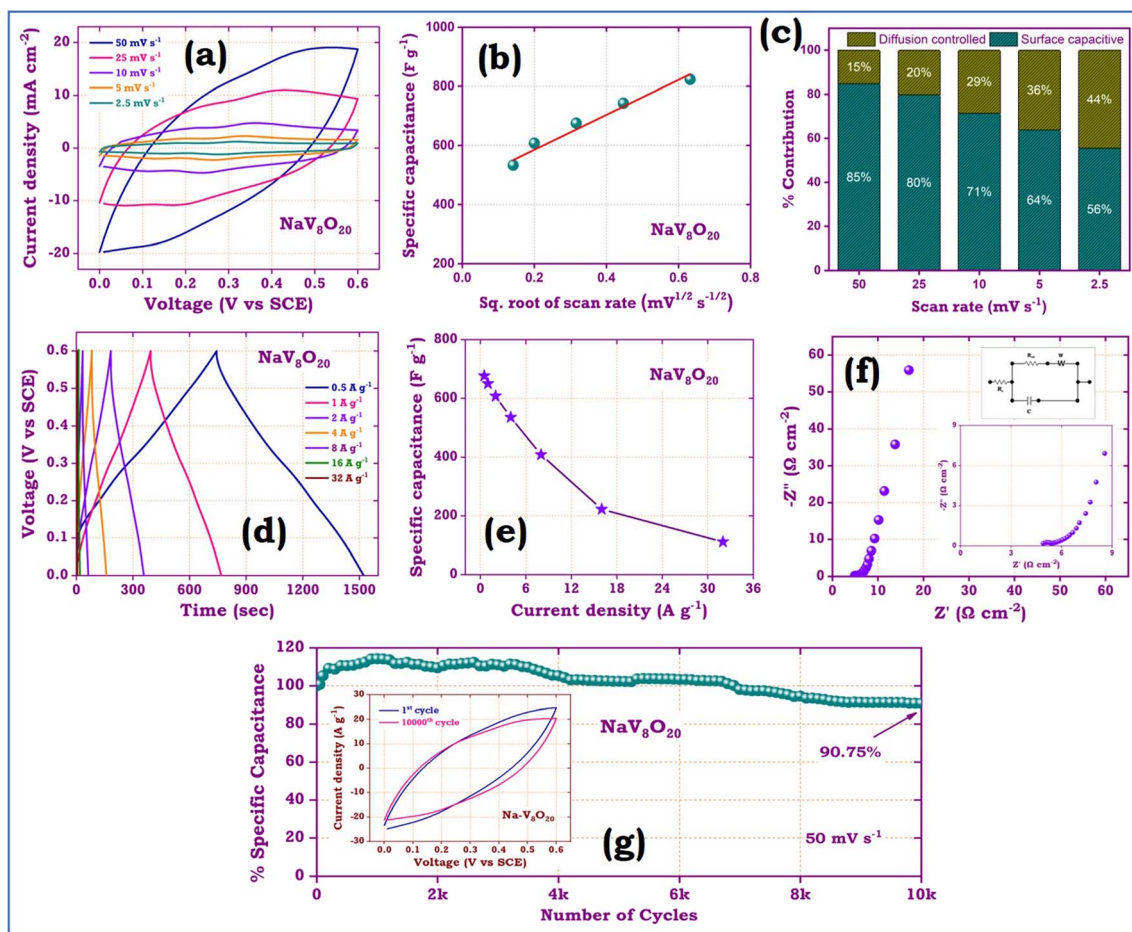


Fig. 5 (a) CV curves, (b)  $C_{sp}$  from CV curves versus square root of scan rate, (c) histogram for charge storage contribution; (d) GCD profile; (e) plot of  $C_{sp}$  vs. current density; (f) Nyquist plots and (g) cyclic stability of NaV<sub>8</sub>O<sub>20</sub> electrode.



During the cathodic CV scans (ESI, Fig. S3(d)†), two reduction peaks at  $\sim 0.2$  and  $\sim 0.35$  V vs. SCE arose because of the reduction of vanadium ionic species from  $V^{5+} \rightarrow V^{4+}$  and  $V^{4+} \rightarrow V^{3+}$ . Whereas, in the anodic CV scans, two oxidation peaks at  $\sim 0.1$  and  $\sim 0.24$  V vs. SCE were detected due to the oxidation of vanadium ionic species from  $V^{3+} \rightarrow V^{4+}$  and  $V^{4+} \rightarrow V^{5+}$ . Interestingly, by increasing the scan rates, the cathodic peaks gradually shifted towards the higher potentials, indicating that rapid and scan rate-dependent redox reactions occurred near the  $NaV_8O_{20}$  electrode surface.<sup>35</sup> Fig. 5d shows the GCD curves of the  $NaV_8O_{20}$  electrode, performed 0.5 to 32 A g<sup>-1</sup> of current densities. All the GCD curves show an ideal triangular and symmetric nature with a slight hump at  $\sim 0.4$  V vs. SCE, which indicates fast faradaic reactions and high reversibility of  $Na^+$  ions on the  $NaV_8O_{20}$  electrode surface. Additionally, the GCD curves show extended charging–discharging characteristics with a negligible  $iR$  drop, signifying the outstanding rate capability. The  $C_{sp}$  values were calculated using eqn (E2) (see ESI†) and summarized in the ESI, Table S3.† The  $NaV_8O_{20}$  electrode displays an optimal  $C_{sp}$  of 676 F g<sup>-1</sup> at a current density of 0.5 A g<sup>-1</sup>; afterwards, it decreases with an increase in current density (as shown in Fig. 5e). At a 32 A g<sup>-1</sup> current density, the  $NaV_8O_{20}$  electrode still provides the admirable  $C_{sp}$  of 142 F g<sup>-1</sup>. We believe these excellent  $C_{sp}$  values mainly arise from the pre-insertion of  $Na^+$  into the  $V_8O_{20}$  structure, which offers a high electroactive surface area, enhances the overall electronic conductivity and simplifies the  $Na^+$  ions insertion/de-insertion process. The electrode material's low electric and electrochemical resistances are typically preferred for the supercapacitor's longer cyclic lifespan and versatile applicability. The electrochemical resistance characteristics of the  $NaV_8O_{20}$  electrode material were evaluated through EIS measurements. The Nyquist plot for the  $NaV_8O_{20}$  electrode is presented in Fig. 5f. In the high-frequency region of the Nyquist plot, two key components are observed: the equivalent series resistance ( $R_s$ ) and the charge transfer resistance ( $R_{ct}$ ).  $R_s$  represents the overall internal (ohmic) resistance, which encompasses the active electrode material's intrinsic resistance, the electrolyte's ionic resistance, and the contact resistance at the electrode/electrolyte interface.<sup>36–38</sup> Meanwhile, the diameter of the semi-circle represents the charge transfer resistance, which is associated with faradaic reactions. The simulated circuit provided that the  $NaV_8O_{20}$  electrode possesses  $R_s = 4.9 \Omega$  and  $R_{ct} = 0.55 \Omega$ . These low resistance values are attributed to the high electronic conductivity, large electroactive surface area, enhanced electron transport rate, and reduced charge transfer resistance of the  $NaV_8O_{20}$  electrode, which collectively enable rapid faradaic reactions at the electrode/electrolyte interface. In the Nyquist plot, the Warburg impedance ( $W$ ) is associated with the vertical line in the low-frequency region, corresponding to the intercalation or diffusion of electrolyte ions inside the active electrode material. The slope of the Warburg impedance approaches a 90° angle, signifying the capacitive behaviour of the  $NaV_8O_{20}$  electrode.

Long cyclic life is one of the significant parameters for electrochemical supercapacitor applications. The electrochemical cycling stability over the  $NaV_8O_{20}$  electrode (Fig. 5g) was

performed at a scan rate of 50 mV s<sup>-1</sup> over 10 000 cycles. In the initial  $\sim 1000$  cycles, the  $C_{sp}$  gradually increases and then tends to stabilize for further CV cycling. This is because of the activation of the  $NaV_8O_{20}$  electrode through the insertion/de-insertion of  $Na^+$  ions and complete infiltration of  $Na^+$  electrolyte ions inside the active electrode material surface.<sup>21</sup> The  $NaV_8O_{20}$  electrode retains 90% of its initial  $C_{sp}$  after 10 000 CV cycles. This outstanding cycling stability is mainly ascribed to the layered and 1D nanobelts structured assembly of the  $NaV_8O_{20}$  electrode, which offers significant redox active sites and efficiently accommodates the possible volume deviations through  $Na^+$  ions insertion/de-insertion process, which ultimately enhances the charge storage capacity of the  $NaV_8O_{20}$  electrode material.

## 4.2 Electrochemical performance of the AC// $NaV_8O_{20}$ asymmetric supercapacitor (ASC) device

To evaluate the practical applications of the  $NaV_8O_{20}$  nanobelts for supercapacitors, the AC// $NaV_8O_{20}$  ASC coin cell device (Fig. 1d) was fabricated. Before the fabrication of the SC device, the  $NaV_8O_{20}$  (cathode) and AC (anode) were electrochemically evaluated in a three-electrode system to assess the appropriate working voltage window.

Fig. 6a shows the CV profile for the AC and  $NaV_8O_{20}$  electrodes carried out in a three-electrode system at a scan rate of 50 mV s<sup>-1</sup> in 1 M  $Na_2SO_4$  electrolyte. The CV curves of  $NaV_8O_{20}$  and AC electrodes are nearly rectangular, measured between the voltage windows of 0 to 0.6 V vs. SCE and  $-1$  to 0 V vs. SCE, respectively. These results imply that integrating the cathode and anode can provide a 1.6 V working voltage. After the fabrication of the ASC device, first, CV measurements with different voltages from 1 to 1.6 V (Fig. 6b) were performed, which indicates the actual working voltage can be extended up to 1.6 V without showing any alteration in CV curves due to aqueous electrolyte redox processes. Fig. 6c displays the CV curves of the AC// $NaV_8O_{20}$  ASC device, measured at 10 to 100 mV s<sup>-1</sup> scan rates. These CV curves retain the quasi-rectangular shape observed for  $NaV_8O_{20}$  nanobelts electrodes in a three-electrode setup, confirming the pseudocapacitive behaviour of  $NaV_8O_{20}$  nanobelt material. Fig. 6d displays the GCD curves of the AC// $NaV_8O_{20}$  ASC device measured at 0.5 to 16 A g<sup>-1</sup> of current densities, demonstrating good symmetric and triangular GCD curves because of the high reversible charge/discharge processes. The  $C_{sp}$  of the AC// $NaV_8O_{20}$  ASC device (ESI, Table S3†) was calculated using eqn (E2) (ESI†). Fig. 6e shows the change in  $C_{sp}$  with current densities. The highest  $C_{sp}$  of 126 F g<sup>-1</sup> was achieved at 0.5 A g<sup>-1</sup> of current density. Energy ( $E_d$ ) and power densities ( $P_d$ ) of the AC// $NaV_8O_{20}$  ASC device were calculated using the following equations:<sup>39</sup>

$$E_d = \frac{C_{sp} \times (\Delta V)^2}{2 \times 3.6} \quad (1)$$

$$P_d = \frac{E}{\Delta t} \times 3600 \quad (2)$$

where,  $C_{sp}$  is the specific capacitance of the ASC device (F g<sup>-1</sup>),  $\Delta V$  is the voltage window, and  $\Delta t$  is the discharging time of the ASC device. Fig. 6f depicts the Ragone plot of the fabricated ASC



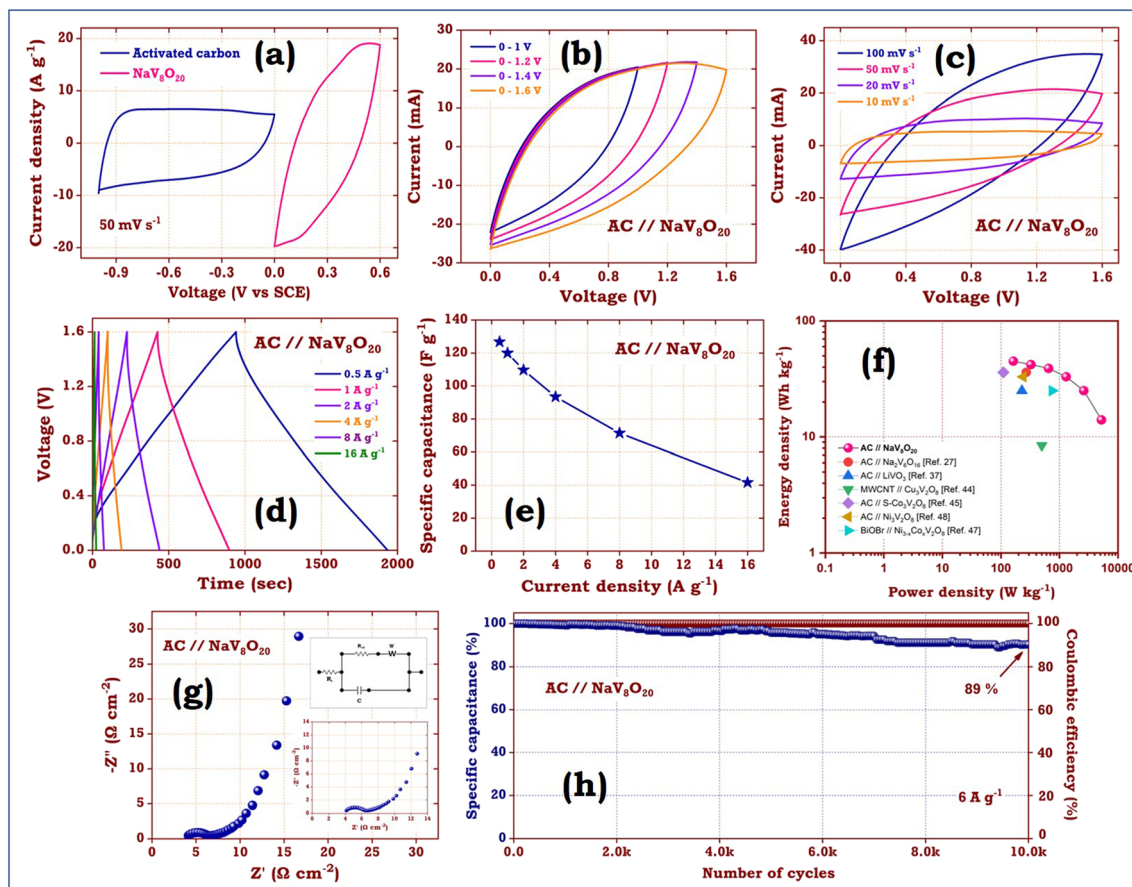


Fig. 6 (a) AC and  $\text{NaV}_8\text{O}_{20}$  electrodes CV curves in the three-electrode system, (b) CV curves for different voltage windows and (c) for different scan rates; (d) GCD curves for different current densities; (e)  $C_{\text{sp}}$  vs. current densities; (f) Ragone plot; (g) Nyquist Plot and (h) cyclic stability of AC// $\text{NaV}_8\text{O}_{20}$  ASC device.

coin cell device. The AC// $\text{NaV}_8\text{O}_{20}$  ASC device shows the highest  $E_d$  of  $45 \text{ W h kg}^{-1}$  at  $P_d$  of  $163.26 \text{ W kg}^{-1}$  and the highest  $P_d$  of  $5224.48 \text{ W kg}^{-1}$  at  $E_d$  of  $14.8 \text{ W h kg}^{-1}$  (ESI, Table S4†), which is higher than reported supercapacitor devices, as shown in Table 1. The resistive characteristics of the AC// $\text{NaV}_8\text{O}_{16}$  ASC device were evaluated by EIS measurement. The Nyquist plot and the equivalent circuit for the ASC device are shown in Fig. 6g. The AC// $\text{NaV}_8\text{O}_{20}$  ASC device displays  $R_s = 3.8$  and  $R_{\text{ct}} = 2.9 \Omega$ . The cyclic life of the SC is considered a key parameter for its long-term applicability in portable electronic devices. Therefore, the cyclic stability of the AC// $\text{NaV}_8\text{O}_{20}$  ASC device was evaluated at a current density of  $6 \text{ A g}^{-1}$  for 10 000 cycles and presented in Fig. 6h. The AC// $\text{NaV}_8\text{O}_{20}$  ASC device displayed 89% of its initial capacitance value with  $\sim 98\%$  coulombic efficiency. This good cyclic stability can be accredited to (1) the proper combination of faradaic and pseudocapacitive electroactive material, (2) the appropriate interlayer spacing of the electrode materials, and (3) the large surface area of both electrode materials.

A detailed comparative analysis of the  $\text{NaV}_8\text{O}_{20}$  nanobelts developed in this study with the various metal vanadate electrodes for SC devices previously reported in the literature is presented in Table 1.

The comparative results suggest that, compared with reported metal vanadate electrode materials, the  $\text{NaV}_8\text{O}_{20}$

nanobelts consistently demonstrate better electrochemical performance. The  $\text{NaV}_8\text{O}_{20}$  nanobelts exhibit significantly higher capacitance with an operating window of 1.6 V, surpassing many vanadium-based alternatives. They offer excellent cyclic stability, retaining their electrochemical performance over 10 000 cycles, making them ideal for high-performance applications. Here, the key contributor to this exceptional electrochemical performance is the use of the  $\text{Na}_2\text{SO}_4$  electrolyte, which provides numerous advantages, such as high ionic conductivity, high ionic mobility, and less corrosive compared to KOH. It also ensures compatibility with the  $\text{Na}^+$  pre-insertion in the electrode material, allowing for efficient ion transfer and enhanced electrochemical reactions. Moreover, the excellent electrochemical performance of the  $\text{NaV}_8\text{O}_{20}$  nanobelts can be primarily ascribed to the pre-insertion of  $\text{Na}^+$  ions during synthesis. This pre-insertion enhances the redox active sites, provides an efficient diffusion path for  $\text{Na}^+$  ions, and ensures optimal interlayer spacing, facilitating effective  $\text{Na}^+$  ion intercalation. Additionally, the high electrical conductivity of the  $\text{NaV}_8\text{O}_{20}$  nanobelts further improves their overall electrochemical behaviour, leading to enhanced charge storage and cyclic stability. These features collectively position  $\text{NaV}_8\text{O}_{20}$  nanobelts as an auspicious material for next-generation energy storage technologies.





Table 1 Comparative study of NaV<sub>8</sub>O<sub>20</sub> electrode and ASC device with previously reported metal vanadate electrodes and devices

Vanadate material	Three-electrode system			Supercapacitor device				
	Method	Morphology	Electrolyte	$C_{sp}$ (F g <sup>-1</sup> @ A g <sup>-1</sup> )	Electrode stability (% @ cycles)	Voltage window (V)	$E_d$ (W h kg <sup>-1</sup> )	$P_d$ (W kg <sup>-1</sup> )
LVO <sub>3</sub>	Chemical route	Nanorods	1 M Li <sub>2</sub> SO <sub>4</sub>	426 @ 0.5	80 @ 2000	1.6	25	3600
Li <sub>3</sub> VO <sub>4</sub> /carbon	Solid-state	Nanofibers	0.1 mol per L (NH <sub>4</sub> ) <sub>2</sub> SO <sub>4</sub>	—	—	3.8	110	3870
β-Na <sub>0.33</sub> V <sub>2</sub> O <sub>5</sub>	Hydrothermal	Nanobelts	1 M LiCO <sub>3</sub> /PC	320 @ 5 mV s <sup>-1</sup>	—	1	47	5000
β-Na <sub>0.33</sub> V <sub>2</sub> O <sub>5</sub>	Spin coating	Nanowires	1 M LiCO <sub>3</sub> /PC	498 @ 0.4	90 @ 1500	—	—	—
NaV <sub>3</sub> O <sub>8</sub>	Electrochemical conversion	Nanosheets	1 M Na <sub>2</sub> SO <sub>4</sub>	640.10 @ 0.5	87.7 @ 2000	1.6	62.6	7200
Na <sub>2</sub> V <sub>6</sub> O <sub>16</sub>	Chemical route	Nanobelts	1 M Na <sub>2</sub> SO <sub>4</sub>	455 @ 0.5	90 @ 5000	1.6	42	4300
Na <sub>6</sub> V <sub>10</sub> O <sub>28</sub>	Chemical route	Microtubes	1 M LiClO <sub>4</sub> + PC	354 @ 0.1	—	2.8	73	6238
KV <sub>5</sub> O <sub>8</sub>	Chemical route	Nanobelts	0.5 M K <sub>2</sub> SO <sub>4</sub>	677 @ 0.5	89 @ 5000	1.8	51	4127
Cu <sub>3</sub> V <sub>2</sub> O <sub>8</sub>	SILAR	Nanopebbles	1.5 M NaClO <sub>4</sub>	443 @ 5 mV s <sup>-1</sup>	85 @ 2000	1.15	8.4	345
S-Co <sub>3</sub> V <sub>2</sub> O <sub>8</sub>	Hydrothermal	Nanosheets	6 M KOH	410 mA h g <sup>-1</sup> at 2 A g <sup>-1</sup>	94 @ 4000	1.5	36	11 600
Mn <sub>2</sub> V <sub>2</sub> O <sub>7</sub>	Hydrothermal	Nanoparticles	3 M KOH	327 C g <sup>-1</sup> at 1 A g <sup>-1</sup>	93 @ 10 000	1.5	44.3	807.4
Ni <sub>3-x</sub> Co <sub>x</sub> V <sub>2</sub> O <sub>8</sub>	Hydrothermal	Oval-shaped	3 M KOH	391 @ 0.5	98 @ 8000	1.3	25.5	7500
Ni <sub>3</sub> V <sub>2</sub> O <sub>8</sub> /Ni	Hydrothermal	Nanosheets	3 M KOH	1300 @ 1	80 @ 7000	1.2	33.2	2400
NiCo <sub>2</sub> V <sub>2</sub> O <sub>8</sub>	Microwave-assisted	Microspheres	2 M KOH	113.2 mA h g <sup>-1</sup> at 0.5 A g <sup>-1</sup>	82 @ 10 000	1.6	34.8	4000
NaV <sub>8</sub> O <sub>20</sub>	Hydrothermal	Nanobelts	1 M Na <sub>2</sub> SO <sub>4</sub>	676 @ 0.5	90 @ 10 000	1.6	45	5224
								<b>89 @ 10 000</b>
								<b>This work</b>

### 4.3 Practical application of the AC//NaV<sub>8</sub>O<sub>20</sub> ASC coin cell device

Fig. 7 displays the AC//NaV<sub>8</sub>O<sub>20</sub> ASC coin cell device for its practical application by glowing the 28-red LED-contained panel. Each red LED bulb requires 1.2 V to shine; therefore, two AC//NaV<sub>8</sub>O<sub>20</sub> coin cell devices were connected in series for lighting. Then, fabricated ASC coin cell devices were first charged up to 1.6 V and discharged through an LED panel. These two devices can easily illuminate LED panels up to ~120 s (ESI Media File†), demonstrating the successful practical applicability of NaV<sub>8</sub>O<sub>20</sub> nanobelts material for SC application.

### 4.4 Electrochemical performance of NaV<sub>8</sub>O<sub>20</sub> nanobelts for Na<sup>+</sup> ion battery

To identify the electrochemical features of the NaV<sub>8</sub>O<sub>20</sub> nanobelts for the Na-ion battery, half-cell devices were assembled using Na metal foil as the reference/counter electrode. Fig. 8 shows the electrochemical measurements of a fabricated Na-ion battery performed in the voltage range of 1.5–4.0 V vs. Na/Na<sup>+</sup> to investigate the electrochemical redox reaction kinetics of the NaV<sub>8</sub>O<sub>20</sub> nanobelts. Fig. 8a shows the initial three CV curves recorded at a scan rate of 0.1 mV s<sup>-1</sup> vs. Na/Na<sup>+</sup>. These CV curves show a significant degree of overlap, particularly during the initial three cycles, indicating excellent reversibility of Na<sup>+</sup> ion insertion/extraction between the electrodes. The CV profiles exhibit two distinct intercalation peaks at 1.997 V and 3.631 V, along with two corresponding de-intercalation peaks at 3.528 V and 1.75 V. These peaks are associated with the reversible intercalation and de-intercalation of Na<sup>+</sup> ions within the NaV<sub>8</sub>O<sub>20</sub> structure, contributing to the material's high-capacity retention. Along with this, some minor oxidation peaks (3.075 and 3.204 V) and reduction (2.783 and 2.991 V) peaks were observed. These peaks may arise due to the multistep process of Na<sup>+</sup> ions insertion and de-insertion.<sup>53,54</sup>

The CV measurements for 0.1 to 1 mV s<sup>-1</sup> scan rates were also performed (ESI, Fig. S3(e)†) to examine the rate performance and reaction kinetics. Fig. S3(e)† indicates that the peak positions slightly shift towards higher potentials, specifying that the Na<sup>+</sup> ions' insertion and de-insertion kinetics are closely related at different scan rates. To evaluate the Na<sup>+</sup> ions diffusion characteristics for the NaV<sub>8</sub>O<sub>20</sub> electrode material, the diffusion coefficient ( $D$ ) is determined by the Randles–Sevcik equation:<sup>55</sup>

$$i_p = 0.4463 \times n \times F \times C \times A \times \left[ \frac{nF\theta D}{RT} \right]^{1/2} \quad (3)$$

where  $i_p$  is the peak current,  $n$  is the electron density,  $C$  is the electrolyte molar concentration,  $F$  is Faraday's constant (C mol<sup>-1</sup>),  $A$  is an electroactive area of an electrode (cm<sup>2</sup>),  $D$  is the diffusion coefficient of a redox mediator (cm<sup>2</sup> s<sup>-1</sup>),  $\theta$  is the scan rate (mV s<sup>-1</sup>),  $R$  is the gas constant (J K<sup>-1</sup> mol<sup>-1</sup>) and  $T$  is the temperature (K).

To calculate  $D$  from the CV measurements, the above formula can be modified as follows:

$$\frac{i_p}{\theta^{1/2}} = 2.69 \times 10^5 \times A \times C \times D^{1/2} \times n^{1/2} \quad (4)$$



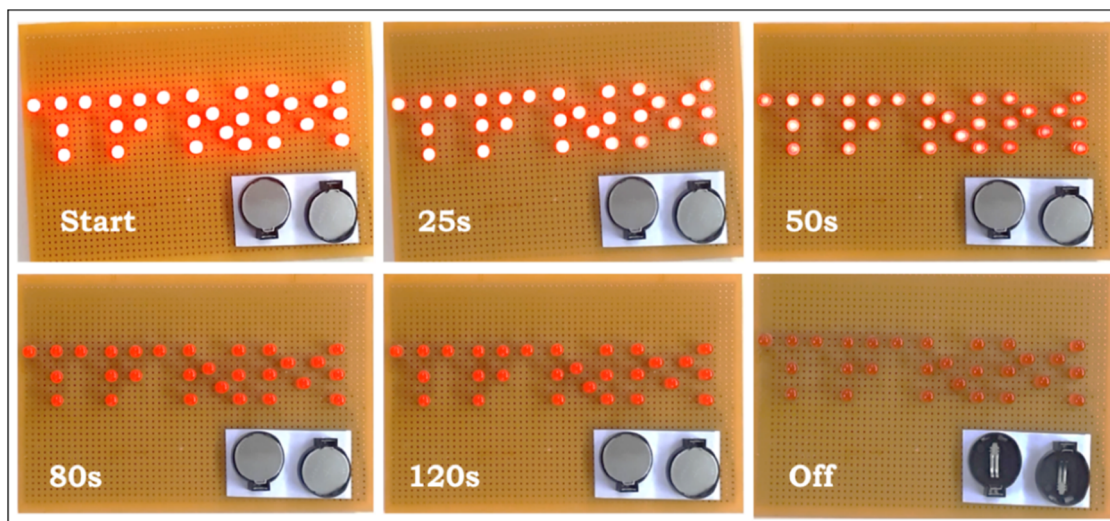


Fig. 7 Device demonstration of the AC//NaV<sub>8</sub>O<sub>20</sub> ASC coin cell device.

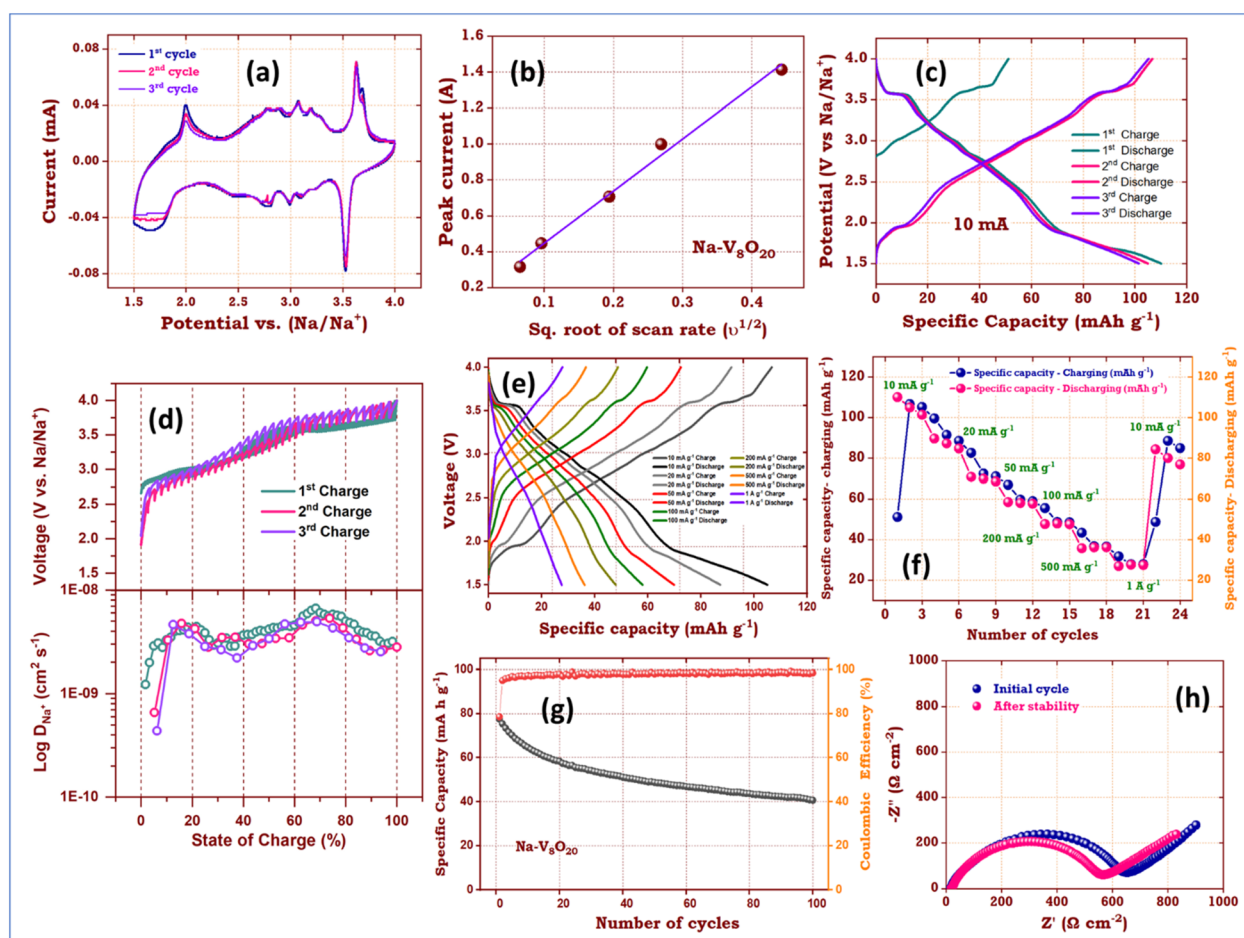


Fig. 8 (a) Initial three CV profiles at  $0.1 \text{ mV s}^{-1}$  between 1.5–4.0 V, (b) peak current vs. square root of the scan rates, (c) galvanostatic charge/discharge curves at  $10 \text{ mA g}^{-1}$ , (d) galvanostatic intermittent titration technique (GITT) profile and corresponding the diffusion coefficient of Na<sup>+</sup> ions and corresponding the diffusion coefficient of sodium ions during 1<sup>st</sup>, 2<sup>nd</sup> and 3<sup>rd</sup> charge at  $100 \text{ mA g}^{-1}$  for 300 s followed by 2 h, (e) GCD curves with different current densities from 10 to  $1000 \text{ mA g}^{-1}$ , (f) rate capability plot, (g) cycling performance at a current density of  $100 \text{ mA g}^{-1}$  and (h) Nyquist plot Na-ion battery device.



Table 2 Summary of diffusion coefficient of various sodium-ion cathode materials reported in the literature with our work

Electrode material	Current density	GITT-pulse duration/relaxation time	Diffusion coefficient ( $D_{\text{Na}^+}$ )	Ref.
NaVPO <sub>4</sub> F	14.3 mA g <sup>-1</sup>	600 s/2 h	10 <sup>-11</sup> to 10 <sup>-10</sup> cm <sup>2</sup> s <sup>-1</sup>	57
Na <sub>3</sub> MnTi(PO <sub>4</sub> ) <sub>3</sub>	11.7 mA g <sup>-1</sup>	10 min/30 min	10 <sup>-10</sup> to 10 <sup>-15</sup> m <sup>2</sup> s <sup>-1</sup>	58
Fe-HCF (iron-based hexacyanoferrate)	—	600 s/2 h	~10 <sup>-11</sup> cm <sup>2</sup> s <sup>-1</sup>	59
NaFePO <sub>4</sub>	3.85 mA g <sup>-1</sup>	1 h/15 h	~10 <sup>-17</sup> cm <sup>2</sup> s <sup>-1</sup>	60
Na <sub>0.66</sub> Ca <sub>0.05</sub> [Ni <sub>0.25</sub> Li <sub>0.11</sub> Mn <sub>0.64</sub> ]O <sub>1.95</sub> F <sub>0.05</sub>	13 mA g <sup>-1</sup>	30 min/5 h	10 <sup>-11</sup> to 10 <sup>-12</sup> cm <sup>2</sup> s <sup>-1</sup>	61
Na <sub>8</sub> Fe <sub>5</sub> (SO <sub>4</sub> ) <sub>9</sub> @rGO	—	—	10 <sup>-11</sup> to 10 <sup>-12</sup> cm <sup>2</sup> s <sup>-1</sup>	62
<b>NaV<sub>8</sub>O<sub>20</sub></b>	<b>100 mA g<sup>-1</sup></b>	<b>300 s/2 h</b>	<b>~10<sup>-9</sup> cm<sup>2</sup> s<sup>-1</sup></b>	<b>This work</b>

where,  $\frac{i_p}{\theta^{1/2}}$  is the slope of the plot of peak current vs. the square root of the scan rate (Fig. 8b). Using this formula, the calculated

Na<sup>+</sup> ions diffusion coefficient value of the NaV<sub>8</sub>O<sub>20</sub> cathode material was found to be  $4.10 \times 10^{-7}$  cm<sup>2</sup> s<sup>-1</sup>, reflecting faster Na<sup>+</sup> ion diffusion inside NaV<sub>8</sub>O<sub>20</sub> cathode material. This excellent diffusion coefficient value results from (1) 1D nanobelts morphology and high surface area of NaV<sub>8</sub>O<sub>20</sub> nanobelts material, which provides a long pathway for easier ion migration, and (2) high electrical conductivity, which is helpful for an efficient charge transfer process. Fig. 8c displays the GCD plot of the NaV<sub>8</sub>O<sub>20</sub> material at a 10 mA g<sup>-1</sup> of current densities for the initial three cycles. Here, the plateaus are well corroborated with the CV curves and possess a maximum reversible discharge capacity of 110 mA h g<sup>-1</sup>. These three GCD curves slightly overlapped, which showed excellent ionic diffusion features and structural stability, complementing the electronic conductivity of the NaV<sub>8</sub>O<sub>20</sub> cathode material. The galvanostatic intermittent titration technique (GITT) was performed to analyze the diffusion kinetics of Na<sup>+</sup> ions. During the measurement, closed-circuit voltage (CCV) was obtained by applying a current pulse of 100 mA g<sup>-1</sup> for 300 s, followed by a relaxation period of 2 h to measure the quasi-open circuit voltage (QOCV). The diffusion coefficient of sodium ions ( $D_{\text{Na}^+}$ ) of NaV<sub>8</sub>O<sub>20</sub> was calculated from the obtained GITT data according to the following formula:

$$D = \frac{4}{\pi\tau} \left( \frac{n_m V_m}{S} \right)^2 \left( \frac{\Delta E_s}{\Delta E_t} \right)^2 \quad (5)$$

where  $\tau$  is the duration of the current pulse (s),  $n_m$  is the number of moles of active material (mol),  $V_m$  is the molar volume of the electrode (cm<sup>3</sup> mol<sup>-1</sup>),  $S$  is the electrolyte/electrode contact area (cm<sup>2</sup>),  $\Delta E_s$  is the steady-state voltage change due to the current pulse, and  $\Delta E_t$  is the voltage change during titration.<sup>56</sup> The obtained  $D_{\text{Na}^+}$  values range from  $4.4 \times 10^{-10}$  cm<sup>2</sup> s<sup>-1</sup> to  $6.6 \times 10^{-9}$  cm<sup>2</sup> s<sup>-1</sup>, suggesting the effective diffusion of Na<sup>+</sup> ions. For comparison, the diffusion coefficients of various sodium ion cathode materials are summarized in Table 2. The comparative table implies that the 1D and nanobelt morphology of the NaV<sub>8</sub>O<sub>20</sub> electrode significantly enhances the diffusion of Na<sup>+</sup> ions compared to the other reported materials, highlighting its potential in high-efficiency sodium-ion batteries.

Fig. 8e displays the GCD behavior of second cycles for different rates from 10 mA g<sup>-1</sup> to 1 A g<sup>-1</sup>. The NaV<sub>8</sub>O<sub>20</sub> cathode

provides a maximum discharge capacity of 105, 87, 70, 58, 48, 36, and 28 mA h g<sup>-1</sup> for the current densities of 10, 20, 50, 100, 200, 500, and 1000 mA g<sup>-1</sup>. Additionally, the rate capability of the NaV<sub>8</sub>O<sub>20</sub> cathode is shown in Fig. 8f; three cycles with each rate were tested, and the nature shows the electrochemical stability at a particular rate, but the capacity decreases with an increase in the rate. As the current density reverts to 10 mA g<sup>-1</sup>, the NaV<sub>8</sub>O<sub>20</sub> cathode recovers ~83.2% (84 mA h g<sup>-1</sup>) of its initial capacity. In addition, the cyclic stability measurement was performed over the fabricated Na<sup>+</sup> ion battery. Fig. 8g displays the cyclic stability of the Na<sup>+</sup> ion battery performed for 100 cycles at a current density of 100 mA g<sup>-1</sup>. The Na<sup>+</sup> ion battery device displays ~100% coulombic efficiency while retaining ~52.4% of the initial capacity. The observed capacity fading can be attributed to the forming of a passivation layer on the electrode surface and a reduction in the number of electroactive sites. Additionally, EIS was employed to analyze the reaction kinetics of the NaV<sub>8</sub>O<sub>20</sub> cathode, both before and after cycling, as shown in Fig. 8h. By comparing both the EIS spectra, it can be found that the  $R_{\text{ct}}$  (quasi-semicircle) for the diffusion of sodium ions around ~650  $\Omega$ , which after cycling gradually decreased to ~562  $\Omega$  and shows an increase in ion and electron transport mobility as cycling is undertaken. This may be due to the decrease in the charge transfer characteristic in terms of lack of the electrolyte ions at the surface of the electrodes, activation, or reduction in active surface area during a prolonged cycling test.

## 5. Conclusions

In summary, we have successfully synthesized one-dimensional (1D) and layered sodium vanadate (NaV<sub>8</sub>O<sub>20</sub>) nanobelts using a simple, cost-effective hydrothermal method and demonstrated their dual application as high-performance electrode materials for SC and SIB devices. This study represents a significant advancement as it is the first to explore the electrochemical potential of NaV<sub>8</sub>O<sub>20</sub> nanobelts in both SCs and SIBs. Compared to previously reported metal vanadate-based electrode materials, the NaV<sub>8</sub>O<sub>20</sub> nanobelts exhibit remarkable electrochemical performance owing to their unique 1D and layered structure. This architecture significantly improves the electrochemical properties, leading to excellent cycling stability and superior rate performance for both SCs and SIBs. For SCs, the NaV<sub>8</sub>O<sub>20</sub> nanobelts exhibit an impressive  $C_{\text{sp}}$  of 676 F g<sup>-1</sup> at 0.5 A g<sup>-1</sup> and remarkable cyclic stability, retaining ~90% of the





capacitance after 10 000 cycles. Moreover, the ASC coin cell device (AC//NaV<sub>8</sub>O<sub>20</sub>) delivers an optimal  $E_d$  of 45 W h kg<sup>-1</sup> and  $P_d$  of 5224.48 W kg<sup>-1</sup>, with 89% cyclic stability after 10 000 cycles, demonstrating its suitability for high-performance energy storage systems. In addition to SCs, the NaV<sub>8</sub>O<sub>20</sub> nanobelts were employed as a cathode material for SIBs, showing a high initial charge/discharge capacity of 110 mA h g<sup>-1</sup> and a retention of 52.4% capacity after 100 cycles at 100 mA g<sup>-1</sup>. The GITT measurements reveal a high diffusion coefficient for Na<sup>+</sup> ions, indicating that the NaV<sub>8</sub>O<sub>20</sub> nanobelts enable efficient ion transport during cycling, which is essential for high-rate performance and long-term stability. These results highlight the potential of NaV<sub>8</sub>O<sub>20</sub> nanobelts for dual electrochemical applications, with a simple synthesis process which is both practical and sustainable.

## Data availability

The data supporting this article have been included as part of the ESI.†

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

The authors express their gratitude to the Alexander von Humboldt Foundation for donating the potentiostat. ASV is thankful to Symbiosis International (Deemed University), Pune for the financial support provided through the Major Research Project (MJRP) grant (2024/4900). SSW acknowledges support from the SERB-TARE project (TAR/2020/000257) funded by the Department of Science and Technology (DST), India. SAS extends appreciation to the DST, Government of India, for the PhD fellowship provided under the DST-INSPIRE scheme. Also, this research was supported by Basic Science Research Program through the National Research Foundation of Korea (NRF) funded by the Ministry of Education (No. RS-2019-NR040081).

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