

Cite this: *Chem. Sci.*, 2026, 17, 791

The versatile world of cumulene chemistry

Abhishek Pareek, ^{†a} Yu Qiu, ^{†b} Matthew A. Johnson, ^b Rik R. Tykwinski ^{*b} and Przemysław Gawel ^{*a}

Despite a history spanning over a century, cumulenes are often relegated to the realm of perceived curiosities rather than practical synthetic intermediates as a result of their high intrinsic reactivity. In this review, we bridge the gap between the known synthetic strategies for these “exotic” molecules and the potential of their reactivity. By surveying both even and odd [n]cumulenes, paying particular attention to the most accessible [3]cumulenes, alongside more limited but instructive examples of [4]- and [5]cumulenes, we demonstrate that these sp-hybridized frameworks offer exceptional synthetic versatility. Indeed, cumulenes can function effectively as nucleophiles, electrophiles, and dienophiles, enabling cyclooligomerization, cycloaddition, organometallic coupling, and other transformations. We describe how these reactions utilize the “naked” sp-hybridized carbon atoms of the cumulene and their substantial internal energy to access structurally diverse products that would otherwise be challenging or even impossible to obtain using more traditional routes. In doing so, we aim to showcase their potential in organic synthesis and highlight the opportunities they present for constructing novel molecular architectures. By reframing cumulenes as valuable synthetic building blocks, rather than mere curiosities, this review hopes to persuade chemists to incorporate these intriguing scaffolds more broadly into modern organic synthesis.

Received 17th June 2025
Accepted 21st November 2025

DOI: 10.1039/d5sc04462f

rsc.li/chemical-science

^aInstitute of Organic Chemistry, Polish Academy of Sciences, Kasprzaka 44/52, 01-224 Warsaw, Poland. E-mail: pgawel@icho.edu.pl; p.gawel@cent.uw.edu.pl^bDepartment of Chemistry, University of Alberta, Edmonton, Alberta, T6G 2G2, Canada. E-mail: tykwinsk@ualberta.ca[†] These authors contributed equally.

Introduction

Cumulenes, distinguished by their ‘naked’ sp-hybridized carbon atoms and high internal energy, not only pose significant synthetic challenges but also yield numerous exciting



Abhishek Pareek

Abhishek Pareek was born in Rajasthan, India. He obtained his PhD in Organic Chemistry from the Central University of Rajasthan under Prof. Srinivasarao Yaragorla, where he worked on calcium-catalyzed regioselective cyclizations of propargyl alcohols. He then carried out postdoctoral research in Poland, first at the University of Warsaw with Prof. Marcin Kalek, investigating the regioselective dearomatization

of pyridinium salts under Morita–Baylis–Hillman conditions. He is currently working at the Institute of Organic Chemistry, Polish Academy of Sciences, with Dr Przemysław Gawel. His research focuses on synthesis of polycyclic heteroaromatic materials with tailored optoelectronic and chiroptical properties for applications in organic electronics and advanced photonic devices.



Yu Qiu

Yu (Kate) Qiu was born in Canton, China. She received her BSc degree in Chemistry from the University of Alberta. She is currently a PhD student in the group of Professor Rik R. Tykwinski at the University of Alberta, working on the synthesis and investigation of cumulenes with a focus on their reactivity and structural properties.



reactions, balancing these challenges with rewarding chemical diversity. As defined by IUPAC, cumulenes are linear chains of carbon atoms joined by three or more consecutive double bonds.¹ The structure of cumulenes, as predicted by van't Hoff in 1875, varies with the carbon chain length (Fig. 1).^{2,3} $[n]$ Cumulenes with an even number of sp-carbon atoms (and thus an odd number of double bonds, with odd n) are

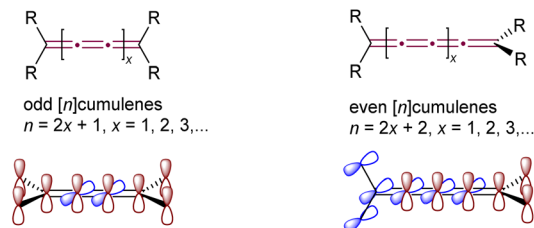


Fig. 1 Classification of cumulenes and schematic depiction of π -orbital systems.



Matthew A. Johnson

Matthew A. Johnson received his PhD degree in 2024 from the University of Alberta under the supervision of Rik Tykwinski. His work focused on the synthesis, reactivity, and elucidation of the electronic structure of $[n]$ cumulenes, with a strong emphasis on even $[n]$ cumulenes.



Rik R. Tykwinski

Rik R. Tykwinski obtained his B.S. at the University of Minnesota–Duluth (1987) and his PhD at the University of Utah (1994). After a PDF at ETH Zürich, he joined the faculty at the University of Alberta (1997). In 2009, he accepted the Chair of Organic Chemistry I at Friedrich-Alexander-Universität Erlangen-Nürnberg in Germany, and he moved back to the University of Alberta in 2016. Physical organic research in the Tykwinski

group centers on synthetic methods for assembling conjugated carbon-rich molecules, characterization of their electronic properties, and materials applications of conjugated systems. Pioneering work on the synthesis of polyynes and $[n]$ cumulenes has led to the discovery of routes to important natural products, as well as ground-breaking discoveries in the study of the one-dimensional carbon allotrope carbyne. We have established a synthetic methodology to form acene derivatives that are vital to elucidating the mechanism and potential of singlet fission/solar energy capture. Our interest in polycyclic aromatic hydrocarbons, coupled with expertise in supramolecular chemistry, is also extended to solve important problems related to the processing of heavy oil (asphaltenes). Tykwinski is the Editor-in-Chief of the *Journal of Physical Organic Chemistry*; he has chaired the Gordon Conference on Physical Organic Chemistry (2017) and recently received the Alfred Bader Award from the Chemical Institute of Canada (2023). In his free time, he can likely be found mountain biking or entertaining his wife and two sons.



Przemysław Gawel

Przemysław Gawel studied Chemistry and Biotechnology at the University of Warsaw, with research experience in H. Hopf's lab at TU Braunschweig. He earned his PhD at ETH Zurich under F. Diederich, working on functional cumulene-based materials, and later joined H. L. Anderson's group at Oxford to study sp-hybridized carbon allotropes. After several years at Cambridge Display Technology Ltd, he returned to Poland in

2021 to establish his independent group at the Institute of Organic Chemistry PAS, which moved to the Centre of New Technologies, University of Warsaw, in 2025. His research focuses on carbon-rich chemistry, optoelectronic materials, and physical organic chemistry.



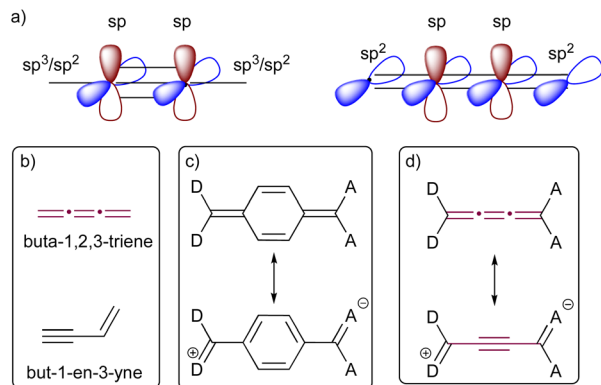


Fig. 2 Proacetylenic character of cumulenes. (a) Comparison of π -orbitals in cumulene and acetylene; (b) structures of buta-1,2,3-triene and but-1-en-3-yne; (c) schematic representation of proaromaticity in quinones; (d) schematic representation of proacetylenic character in cumulenes.^{29–31}

Early interest in $[n]$ cumulenes was driven by their unique structure and potential applications as carbon-rich materials.^{4–7} The initial synthesis of cumulenes, *i.e.*, tetraaryl[3]cumulenes, was achieved by Brand in the early 1920s through HCl elimination from 2,3-dichloro-but-2-enes.^{8,9} The relationship of cumulenes to sp -hybridized carbon allotropes, particularly carbyne and cyclo-carbons, gained equal significance over time.^{10–13} This stemmed from $[n]$ cumulenes, together with polyynes, representing the two fundamental boundary structures of these allotropes, highlighting their role in understanding the nature of carbyne.^{10–12} Significant progress in synthetic methodologies toward carbyne has led to the development of record-breaking [9]cumulenes and a 68-carbon atom-long polyyne,^{14–18} along with cyclo-carbons that exhibited a cumulenic structure in their smaller, aromatic counterparts.^{13,19–24} These remarkable achievements have been instrumental in fostering our understanding of cumulenes.

The latter half of the 20th century saw a significant expansion in cumulenic structures and synthetic methods, which have been reviewed elsewhere.²⁵ In contrast, the rich and diverse reactivity of $[n]$ cumulenes remains dispersed throughout the literature, with no comprehensive summary available. Conversely, polyynes, similar to $[n]$ cumulenes, also contain ‘naked carbon atoms’ with relatively high internal energy²⁶ and are well recognized as valuable building blocks in organic synthesis.^{27,28}

The bonding of sp -hybridized carbon atoms in odd $[n]$ cumulenes shows an alternating pattern of short and long bonds, and as a result, they exhibit electronic characteristics remarkably similar to those of polyynes (Fig. 2). This proacetylenic character of cumulenes is manifested in their structural composition, intrinsic properties, and most notably, their distinctive reactivity patterns.²⁹ This concept is somewhat analogous to the proaromaticity observed in quinoidal chromophores, in which aromatic ring stabilization is the driving force for charge separation (Fig. 2).³⁰ Topological resonance energy (TRE) theory, comprehensively elucidated by Chauvin and coworkers, offers insight into this phenomenon.³¹ According to TRE theory, but-1-en-3-yne is more stable in comparison to its constitutional cumulenic isomer, buta-1,2,3-triene (Fig. 2).³¹ The degree to

which the central C=C bond of a [3]cumulene unit resembles a triple bond is dependent on the electronic character of terminal groups. The polarization induced by substituting a cumulene with electron-donating (D) and electron-accepting (A) groups at the terminal positions leads to a pronounced bond-length alternation (BLA, the difference in bond lengths between the two most central bonds of an odd cumulene).^{32–34} Likewise, formal replacement of alkyl endgroups of an odd $[n]$ cumulene with aryl groups leads to an increased BLA due to increased conjugation.¹⁵

In this review, we aim to bridge the gap between $[n]$ cumulenes as structural curiosities and useful building blocks toward convincing the organic chemistry community that $[n]$ cumulenes offer access to chemical space that is impossible with other types of transformations. We focus in particular on the reactivity of the most common [3]cumulenes while also covering examples involving the more rarely encountered [4]- and [5]cumulenes. Due to their unique structure, $[n]$ cumulenes exhibit remarkable synthetic versatility, participating in reactions as nucleophiles, electrophiles, and dienophiles. They engage in a wide array of transformations, including oxidation, hydrogenation, halogenation, cycloaddition, metalation, metal complexation, and sulfurization. This review presents a comprehensive account of the reactivity of $[n]$ cumulenes, beginning with a historical overview and followed by dedicated sections on cyclooligomerizations, cycloadditions, organometallic transformations, and other relevant reaction types. Given the distinct reactivity patterns of odd- and even-numbered cumulenes, these classes are discussed separately.

Odd cumulenes

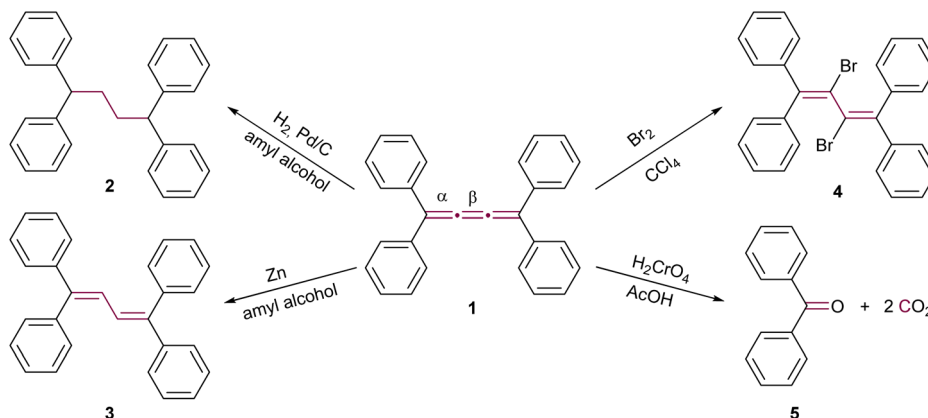
Early cumulene reactions

The initial reactions involving cumulenes, primarily hydrogenations and oxidations, were demonstrated by Brand alongside the first report of the synthesis of an $[n]$ cumulene (Scheme 1).^{8,9} The hydrogenation of tetraphenyl[3]cumulene **1** in the presence of palladium on carbon leads to the complete saturation of the [3]cumulene moiety, yielding 1,1,4,4-tetraphenylbutane **2**. A milder reduction employing zinc gives 1,1,4,4-tetraphenylbuta-1,3-diene **3** (Scheme 1). Electrophilic addition reactions with halogens (Br_2 , I_2 , *etc.*) predominantly occur at the central bond (β) of the [3]cumulene, forming 2,3-dihalobutadienes such as **4**.^{33,35} Notably, bromination may also target a terminal double bond (α) or instigate a domino reaction, leading to the formation of cyclic structures (*vide infra*).^{36–38} Furthermore, oxidation of **1** by chromic acid results in the production of benzophenone **5** and carbon dioxide (Scheme 1). These early 20th-century reactions are typical for non-aromatic unsaturated hydrocarbons, and thus provide limited insight into the distinctive reactivity of cumulenes. Subsequent research yielded numerous instances of analogous reactions, which will not be extensively covered in the discussions that follow.^{36,37,39–41}

Cyclooligomerizations

The reactivity of cumulenes is exemplified through their rich oligomerization chemistry. These unique compounds are





Scheme 1 Early reactions of [3]cumulenes, as reported by Brand in 1921, using the example of tetraphenyl[3]cumulene **1**.^{8,9,33,35–38}

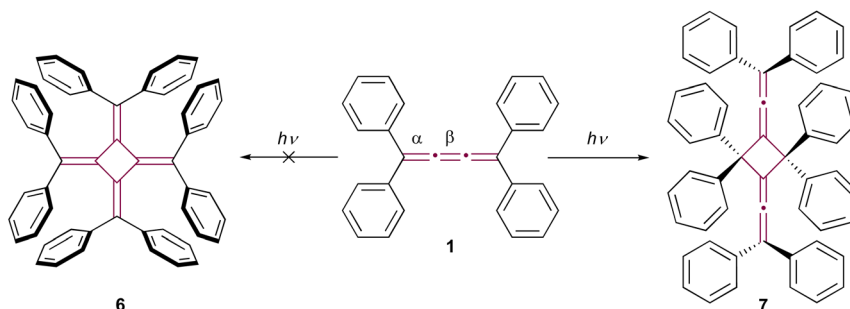
capable of reacting through various mechanisms, including thermal, photochemical, and catalyzed processes. This section will delve into the array of species formed as a result of oligomerization, illustrating the complexity and versatility of these reactions. Remarkably, oligomerization is postulated as a key mechanism in the highly exothermic, and sometimes explosive, decomposition of carbyne – an sp-hybridized allotrope of carbon.⁴²

In his pioneering research, Brand noted that the bright yellow color of tetraphenyl[3]cumulene **1** gradually faded when exposed to sunlight, although the structure of the resultant photodegradation product remained unidentified at the time.^{8,9} It wasn't until 1962 that the structure of the product was proposed to be octaphenyl[4]radialene **6** by Uhler *et al.* (Scheme 2).⁴³ However, this proposal was later refuted by Leiserowitz and colleagues.⁴⁴ They demonstrated that the photochemical dimerization occurred between the terminal double bonds, α , rather than the central β -bond, leading to the head-to-tail dimer **7** (Scheme 2). In the latter studies, the structural elucidation was supported by single-crystal X-ray diffraction analysis and further bolstered by the structural determination of its ozonolysis products.

The early reports of cumulene reactivity were followed by the development of further dimerization methodologies. Heinrich and Roedig reported the thermal dimerization of tetrachloro [3]cumulene **8**, leading to the formation of perchloro[4]radialene **9** (Scheme 3a).⁴⁵ Koster and West explored the thermal

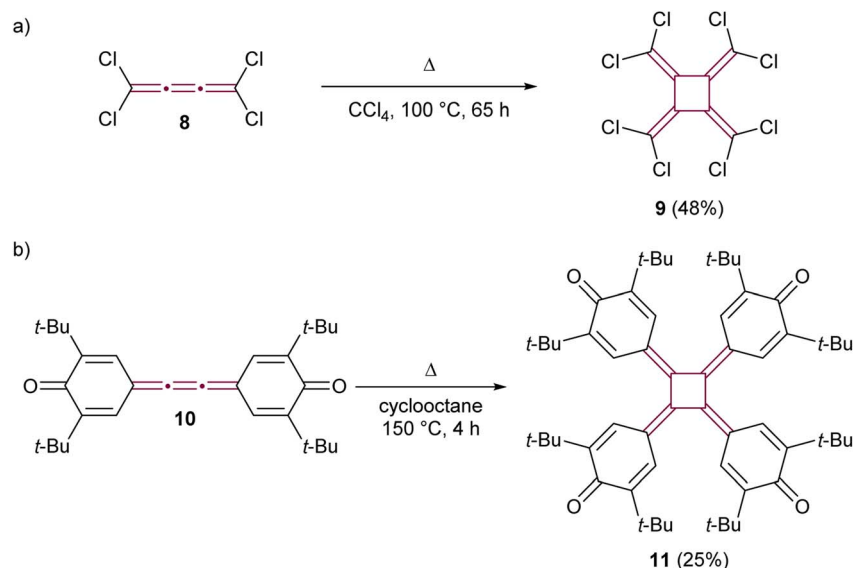
dimerization of extended quinone **10**, successfully synthesizing tetraquinocyclobutane **11**, a compound notable for its unique redox-active properties and the photochemistry of its dianions (Scheme 3b).^{46,47} While these reactions gave moderate yields, the thermal dimerization conditions they pioneered were subsequently adapted for the synthesis of various [4]radialenes from related [3]cumulenes.^{48,49}

The utilization of transition-metal complexes as catalysts marked a significant advancement in cumulene dimerization reactions. In initial experiments, d¹⁰ late transition metal complexes including Pd(0), Pt(0), Rh(I), and Ni(0) were employed to catalyze the dimerization of [3]cumulene **10** into radialene **11** (not shown).⁵⁰ This approach yielded considerable improvements in reaction efficiency over non-catalyzed thermal methods, with the exception of Rh-based catalysts. The complex [Ni(PPh₃)₂(CO)₂] was found to be the most efficient catalyst, delivering 95% isolated yield of radialene after heating for four hours at 80 °C in benzene. Further investigations by Iyoda and coworkers into the cyclooligomerization of tetramethyl[3]cumulene **12**, employing *in situ* generated [Ni(PPh₃)₄], revealed a solvent-dependent outcome.^{51,52} When performed in benzene at 50 °C, the reaction gave [4]radialene **13** (7%) and cyclic dendralene **14** (19%), while [6]radialene **15** was not observed (Scheme 4). Conversely, the same reaction performed in DMF at 50 °C produced [6]radialene **15** in 24% yield, alongside isomer **14** and [4]radialene **13** in 21% and 4% yields, respectively (Scheme 4).



Scheme 2 Photodimerization of tetraphenyl[3]cumulene **1**. Radialene **6** (left) is initially postulated as the product. The product is later revised and verified by X-ray diffraction analysis as **7** (right).^{43,44}



Scheme 3 Thermal cyclodimerization of (a) perchloro[3]cumulene **8** and (b) quinone derivative **10**.^{45–47}

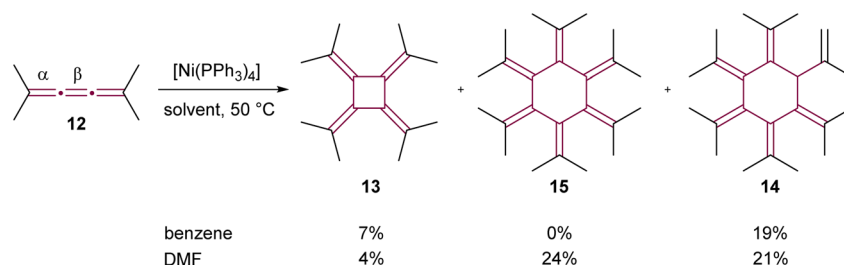
In the proposed reaction mechanism (Scheme 5), initiation of the catalytic cycle occurred through complexation of Ni(0) with the β -bond of **12** to form an η^2 -[3]cumulene nickel complex **A**.⁵¹ Previously, analogous stable complexes with Rh(I) as the central metal have been reported.^{50,53} The process progressed with the addition of a second molecule of **12**, leading to complexation of Ni to the β bond to form a five-membered nickelacycle **B**. These complexes were isolated and characterized in detail in studies conducted by Wilke and Stehling, employing a more stable bipyridine Ni(0) complex.⁵⁴ At this stage, the insertion of a third molecule of **12** into the metallacycle was possible, resulting in the formation of a seven-membered ring **C**. The reductive elimination of these metallacycle complexes subsequently yielded [4]- and [6]radialenes **13** and **15**, respectively (Scheme 5).

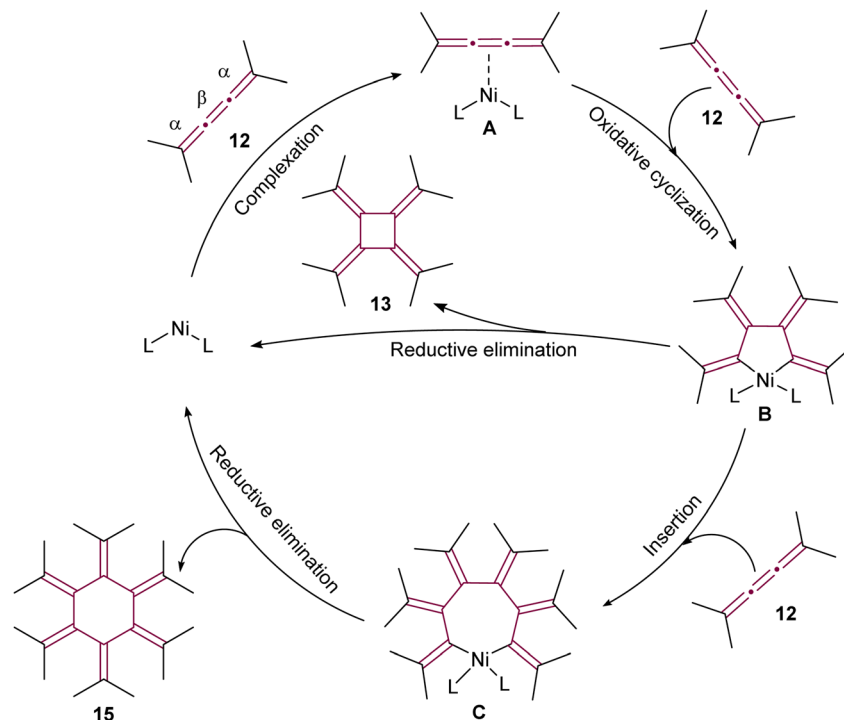
The dimerization of cumulenes has proved to be a good strategy for the synthesis of chromophores featuring cross-conjugated carbon structures, specifically [4]radialenes. Diederich and coworkers have explored novel carbon allotropes employing this strategy. Their initial approach involved attempting to dimerize tetraethynyl[3]cumulenes **16** with the aim of synthesizing perethynylated[4]radialenes **17** (Scheme 6a).⁵⁵ Despite their efforts, neither thermal nor the triphenylphosphine-catalyzed method of West⁵⁰ yielded the anticipated product. Rh-catalyzed methods were also investigated, but they solely resulted in the formation

of a stable η^2 -Rh complex, without yielding the targeted radialene. Later, the same group successfully showed the dimerization of a highly polar push-pull [3]cumulene **18**, employing a $[\text{Ni}(\text{PPh}_3)_2(\text{CO})_2]$ catalyst (Scheme 6b).⁵⁶ The resultant [4]radialene **19** demonstrated intense and broad absorption in the visible spectrum, suggesting its potential suitability for optoelectronic applications.

The dimerization was extended to include longer $[n]$ cumulenes, employing both thermal and Ni-catalyzed techniques. The pioneering work in thermal dimerization of a [5]cumulene was reported by Hartzler in 1966.^{57,58} He observed that tetra(*tert*-butyl)[5]cumulene **20** resolidified after melting at 189 °C forming a white solid, assigned as radialene **21**, which formed through dimerization at the central cumulenic double bond (γ) (Scheme 7a). A few years later, Scott and coworkers reported that tetramethyl[5]cumulene **22** undergoes thermal dimerization through the terminal carbon atoms of the cumulene moiety, resulting in the formation of cyclyne **23** (Scheme 7b).^{59–61} A similar reaction was observed during attempts to synthesize unfunctionalized [3]cumulene, which spontaneously dimerized to cycloocta-1,5-diyne (not shown).^{35,62–64} Hopf and Stang proposed, independently, that [5]cumulenes are intermediates in the synthesis of other polyynic macrocycles.^{65–67}

The thermal dimerization of [5]cumulenes substituted with sterically demanding alkyl groups consistently yielded D_4 -

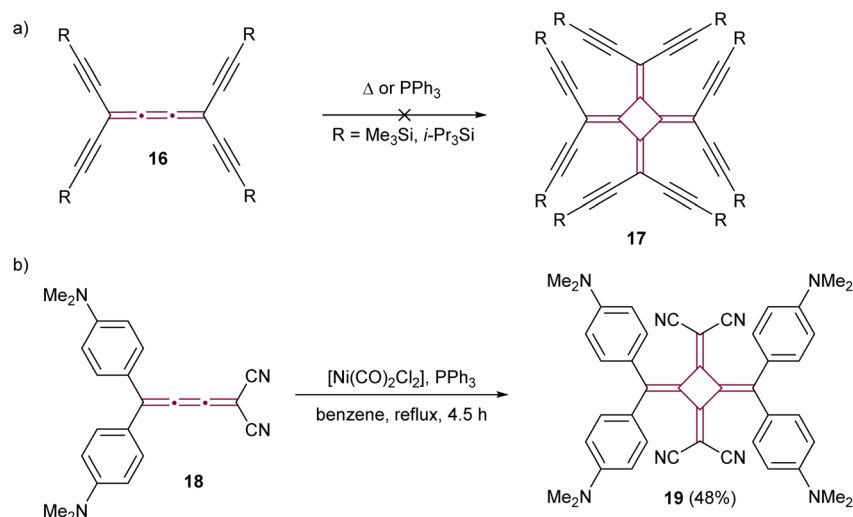
Scheme 4 Ni(0)-catalyzed cyclooligomerization of tetramethyl[3]cumulene **12**.^{51,52}



Scheme 5 Proposed mechanism of Ni(0)-catalyzed cyclooligomerization of tetramethyl[3]cumulene **12** (L = ligand).⁵¹

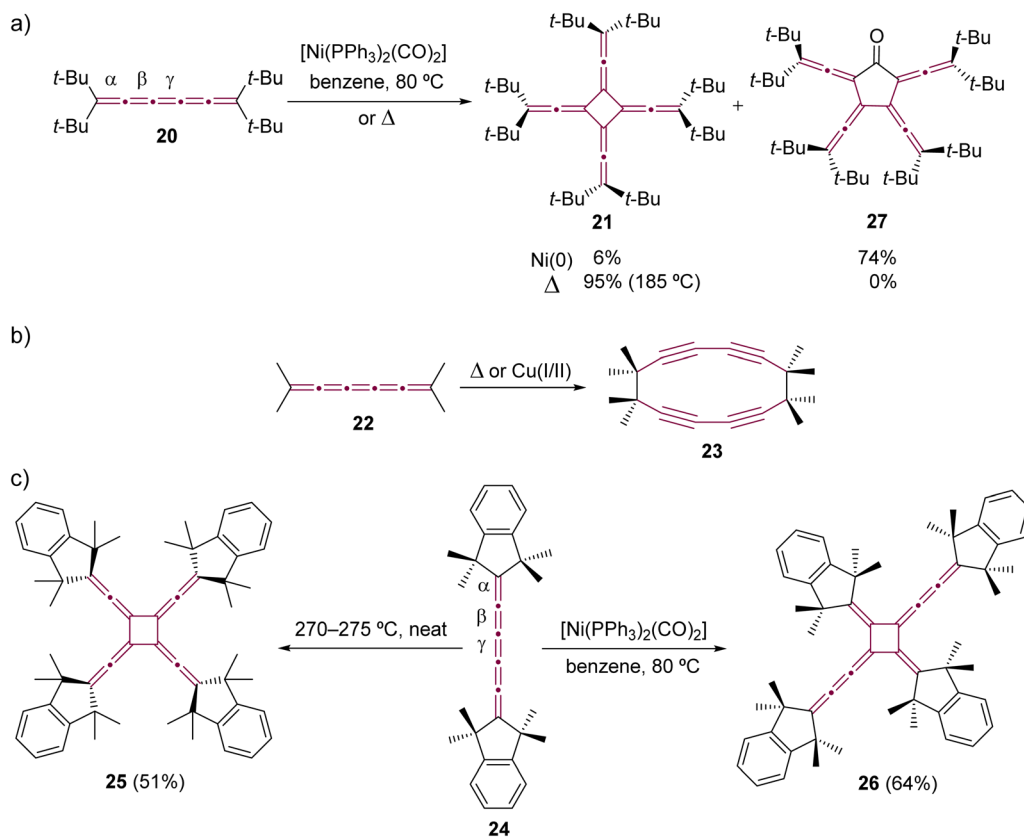
symmetric [4]radialenes, which was later demonstrated in extensive studies by Iyoda and coworkers (Scheme 7c).^{68–71} This research highlights a contrast between the thermal and Ni-catalyzed reactions. The thermal processes predominantly occur at the central γ -bond of the [5]cumulene moiety resulting in D_4 -symmetric [4]radialenes, while the Ni-catalyzed reactions tend to involve the β -bond. This distinction is exemplified in a direct comparison using tetramethylindane endcapped cumulene **24**, which yields two isomeric radialenes: fourfold-symmetric **25** from neat thermal transformation and head-to-tail dimer **26** from a Ni-catalyzed process.⁷⁰ On the other hand, this trend is not always maintained, and cumulenes

decorated with bulky endgroups can yield the same extended [4]radialene *via* either thermal or metal-catalyzed dimerization processes.⁷² For instance, the dimerization of cumulene **20** consistently results in the formation of radialene **21**, irrespective of the method employed (Scheme 7a).^{57,58,69,70} This observed regiochemistry could be attributed to the subtly increased steric hindrance induced by the bulky *tert*-butyl substituents, which obstruct access to α - and β -bonds, guiding the reaction toward the γ -bond. Iyoda and coworkers also observed the formation of [5]radialenone **27** as a byproduct in the Ni-catalyzed reaction (Scheme 7a), which arose from the insertion of carbon monoxide, originating from the



Scheme 6 The attempted synthesis of (a) octaethynyl[4]radialene **17** and (b) the successful synthesis of the push–pull [4]radialene **19**.^{55,56}





Scheme 7 Oligomerizations of [5]cumulenes: (a) tetra(*tert*-butyl)[5]cumulene **20** to extended [4]radialene **21** and cyclic ketone **27**,^{57,58,69,70} (b) tetramethyl[5]cumulene **22** to cyclyne **23**,^{59–61} and (c) comparison of thermal and Ni-catalyzed dimerization reactions of [5]cumulene **24**.^{69,71}

coordination sphere of nickel, into the intermediate nickela-cycle (see Scheme 5).^{69,71}

In comparison to [5]cumulenes substituted with alkyl groups, those functionalized with aromatic substituents, such as phenyl groups, exhibit notably different behavior in both thermal and Ni-catalyzed dimerization processes. For instance, when heated in toluene at 110 °C, tetraphenyl[5]cumulene **28** undergoes a tandem trimerization reaction, leading to the formation of the tricyclodecadiene derivative **29**, as reported by Kawamura and coworkers (Scheme 8a).⁷³ The authors suggested that this reaction initiates with the dimerization of tetraphenyl[5]cumulene **28** into a D_4 -symmetric radialene, analogous to the transformation **20** \rightarrow **21**. This is followed by a [4 + 2] cycloaddition with a third molecule of [5]cumulene **28**, and a subsequent electrocyclic cyclization culminates in the tricyclic product **29**. In contrast, the Ni-catalyzed reaction of **28**, using $[\text{Ni}(\text{PPh}_3)_2(\text{CO})_2]$ in benzene, predominantly yields the D_2 -symmetric head-to-head isomer **30**. This is a notable deviation from the head-to-tail isomer produced in the reaction with compound **24** (Scheme 7c).⁶⁸ It is not clear if this reactivity difference stems from the steric repulsion of tetramethylindane groups in **24** or the π - π stacking interactions between the phenyl rings in **28**, or a combination of both that influence the orientation of cumulenes coordinated at the nickel center during the catalytic cycle. Additional studies reveal that, when catalyzed by $\text{Ni}(\text{PPh}_3)_4$ instead of $[\text{Ni}(\text{PPh}_3)_2(\text{CO})_2]$, the dimerization of **28** produces a minor amount of hexa-1,5-dien-3-yne as a by-product (not shown).⁶⁸ These results suggest that the

reactivity of tetraphenyl[5]cumulene **28** likely involves competing pathways beyond the straightforward head-to-head dimerization.

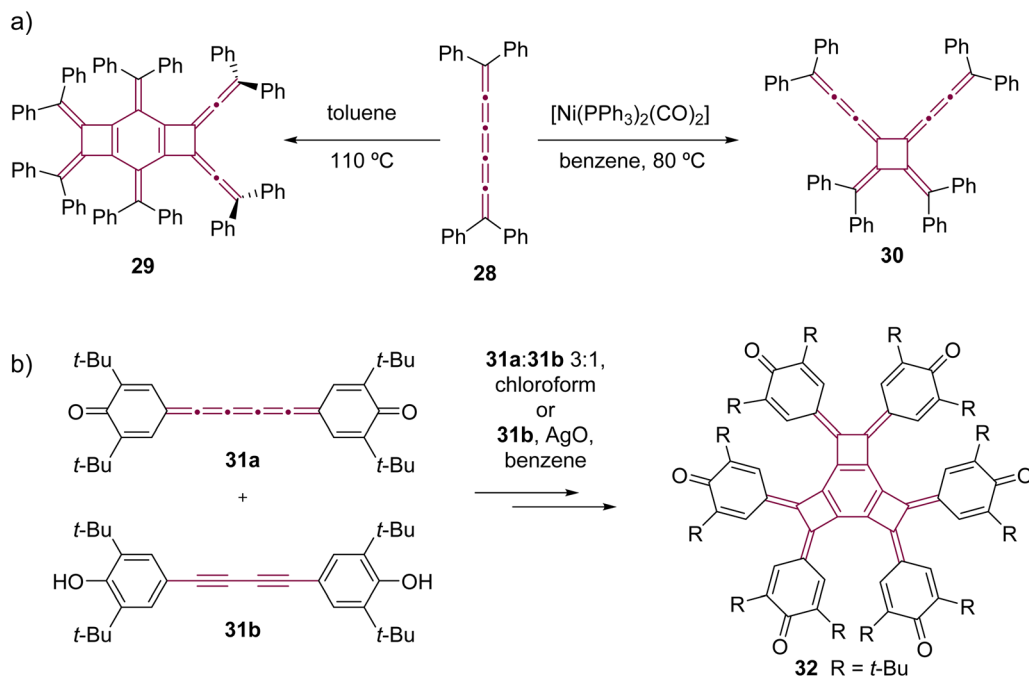
Kawase *et al.* observed that quinocumulene **31a** undergoes trimerization, forming a tricyclobutabenzene derivative **32** (Scheme 8b).⁷⁴ They suggested a mechanism in which the hydroquinone precursor **31b** first oxidizes into a radical intermediate. This radical then initiates dimerization, followed by the addition of [5]cumulene **31a**. The reaction sequence concludes with additional oxidation and cyclization, ultimately yielding the tricyclobutabenzene product **32**.

Only two examples of dimerization involving [n]cumulenes with $n > 5$ have been reported, which was described by Wendiger *et al.*, both leading to cross-conjugated expanded radialenes.⁷⁵ The size of the central radialene ring is determined by the length of the cumulene backbone used as the substrate. A consistent structural feature of the resulting products is the presence of diarylvinyldiene (allene) units located at the vertices of the expanded radialene framework, indicating preference to react at C3 and C3' positions. Hence, tetraaryl[7]- and [9]cumulenes (**33** and **34**, respectively) dimerize to form expanded radialenes **35** and **36** (Scheme 9).

Cycloaddition reactions

To the best of our knowledge, the first report of cycloaddition involving a cumulene features tetra(*tert*-butyl)[5]cumulene **20** (the same report that discussed the dimerization of **20** by



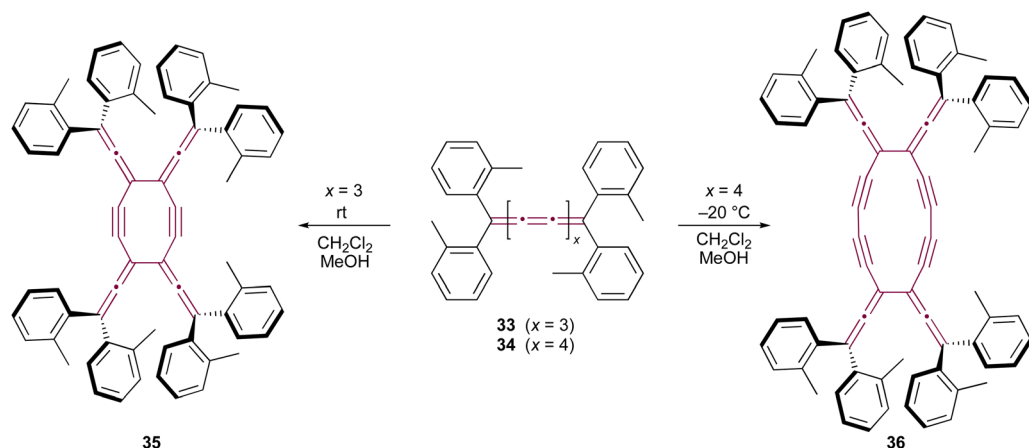


Scheme 8 Oligomerizations of [5]cumulenes: (a) thermal trimerization and Ni-catalyzed dimerization reactions of tetraphenyl[5]cumulene **28**^{68,73} and (b) trimerization of quinocumulene **31a** induced by **31b**.⁷⁴

Hartzler, Scheme 7a).^{57,58} In this study, tetrafluoroethylene reacts with the central γ -bond of cumulene **20** at 200 °C sealed in a Carius tube, forming fluorinated cyclobutene **37** as a result of a formal [2 + 2] cycloaddition reaction (Scheme 10a). Interestingly, the reaction of **20** with ethylene proves to be reversible. The authors postulated the formation of a transient intermediate **38** upon heating with ethylene at 200 °C under pressure, which is found to release ethylene and restore [5]cumulene **20** upon cooling. However, recall that at 200 °C without ethylene, **20** dimerizes to form [4]radialene **21** (Scheme 7a). In the same study, compound **20** is reported to react with hexafluorobut-2-yne at 200 °C, yielding cyclobutene **39** (Scheme 10b). However, under these harsh conditions, **39** isomerizes *via* retro-electrocyclization to form bis[3]cumulene **40**.

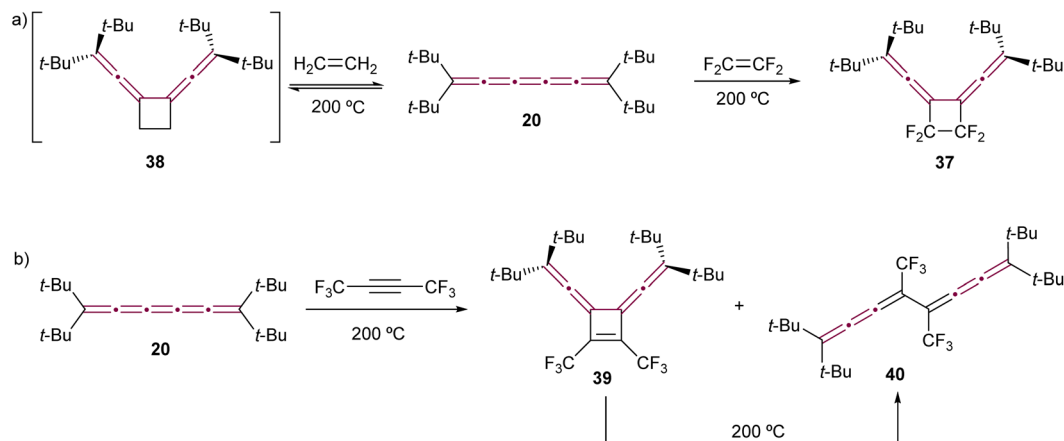
The Diels–Alder reaction, featuring tetraaryl[3]cumulenes as dienophiles, was reported by Ried and Neidhardt.⁷⁶ The reaction of tetraphenyl[3]cumulene **1** and its fluorenylidene end-capped counterpart **41** with various substituted cyclopentadienones **42** produces cycloadducts **43** (Scheme 11). These products are particularly interesting due to their potential conversion, *via* decarbonylation, into *o*-quinodimethanes, compounds that are desirable intermediates in organic synthesis.^{77–79}

In their creative approach, Goroff and coworkers utilized the reversible character of the Diels–Alder reaction with furan for the protection of tetrabromo[3]cumulene **44** (Scheme 12a).⁸⁰ This approach served as a preparatory step for subsequent functionalization of the terminal bromine atoms, specifically through Pd-catalyzed cross-coupling reactions that would



Scheme 9 Thermal dimerization of tetraaryl[7]- **33** and [9]cumulenes **34**.⁷⁵

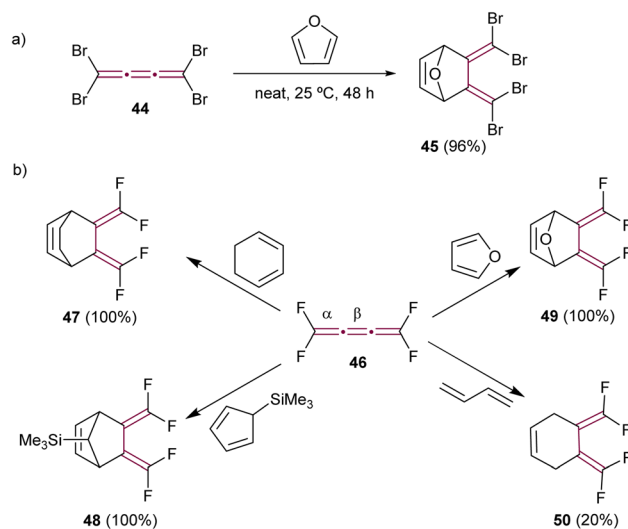




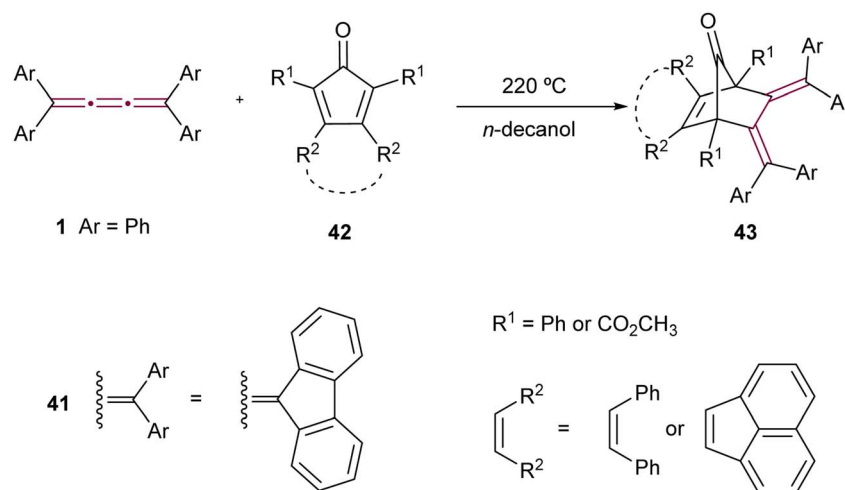
Scheme 10 First reported cycloaddition reactions of tetra(*tert*-butyl)[5]cumulene **20** with (a) ethylene (product not isolated) and tetrafluoroethylene and (b) hexafluorobut-2-yne.^{57,58}

otherwise engage the cumulene core directly. The reaction was carried out neat in furan and produced adduct **45** in a good yield. In contrast, tetrafluoro[3]cumulene **46**, first reported by Martin and Sharkey in 1959, exhibited exceptional reactivity, leading to explosive decomposition in its liquid state.⁸¹ Consequently, it requires storage at extremely low temperatures, *i.e.*, at -196 °C. Nevertheless, this inherent instability did not prevent Ehm and Lentz from reporting on its reactivity in Diels–Alder reactions. They described the reaction of **46** with various dienes, resulting in stable cycloaddition products (selected examples **47**–**50** in Scheme 12b). Notably, these reactions involved the dienes reacting selectively with the central β -bond of tetrafluoro[3]cumulene **46** in all cases.⁸²

Tetracyanoethene (TCNE, **51**) shows notable reactivity with electron-rich acetylenes, which, akin to cumulenes, are characterized by linearly aligned sp -hybridized carbon atoms (Fig. 2). The reaction between TCNE and alkynes typically follows a cycloaddition–retroelectrocyclization (CA–RE) reaction cascade.^{83,84} This transformation proves to be an exceptionally

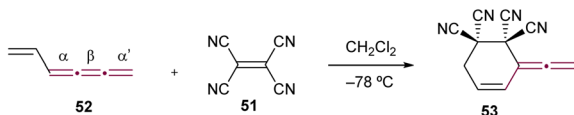


Scheme 12 Diels–Alder reactions of (a) tetrabromo- **44** and (b) tetrafluoro[3]cumulenes **46**.^{80–82}



Scheme 11 Diels–Alder reactions of tetraaryl[3]cumulenes **1** and **41**.⁷⁶



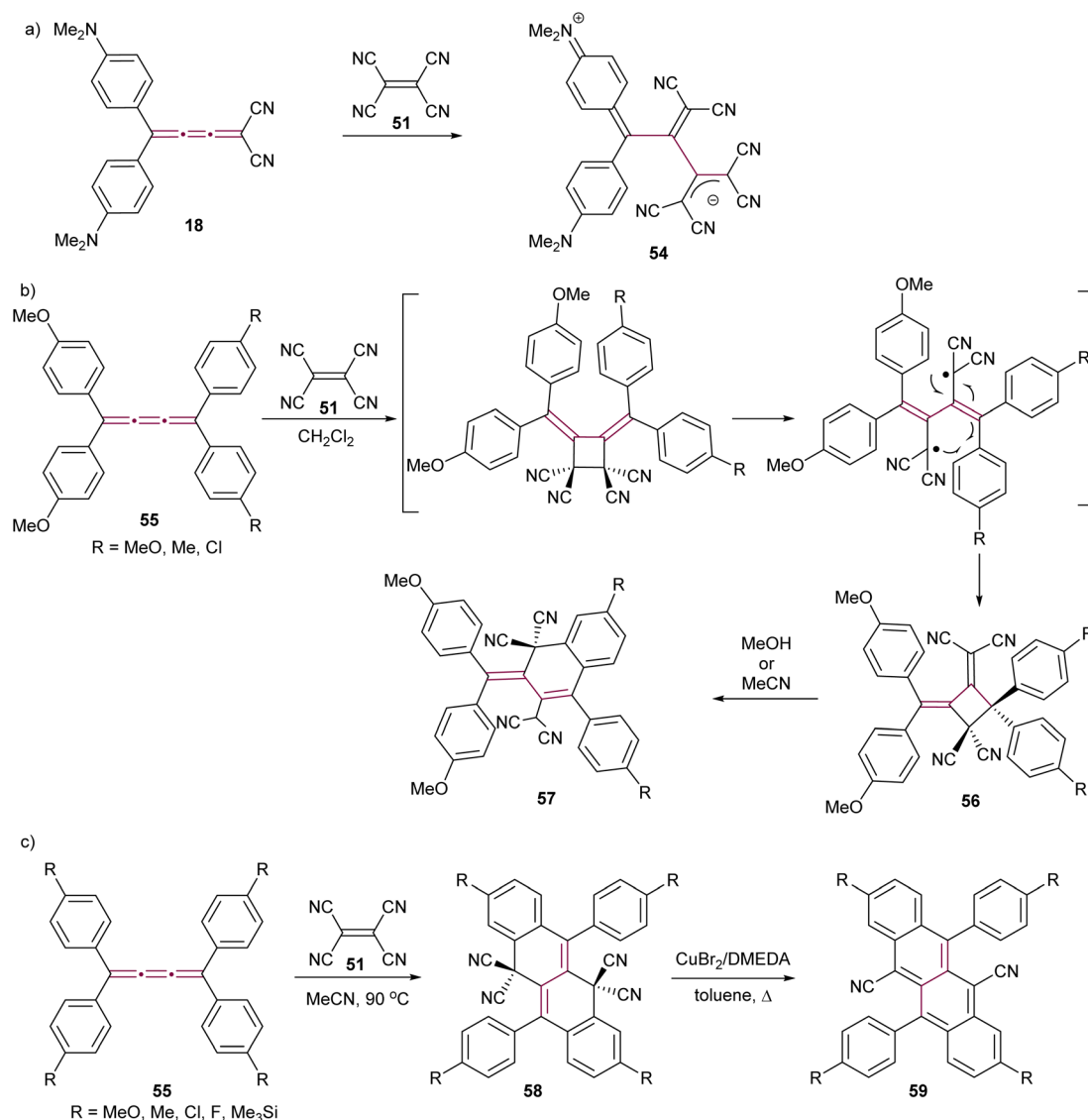


Scheme 13 Diels–Alder reaction of 1,2,3,5-hexatriene **52** with TCNE **51**.⁹⁶

efficient and valuable approach for synthesizing push–pull chromophores with low HOMO–LUMO gaps.^{30,85–95} The analogous reactions between cumulenes and TCNE **51** lead to the formation of a diverse array of products, highlighting their versatility as building blocks in organic synthesis. The initial reaction of this type, reported by Hopf and Maurer, is used to confirm the structure of 1,2,3,5-hexatriene **52**,⁹⁶ which represents one of the twelve possible linear isomers of benzene.⁹⁷ In this reaction, cumulene **52** acts as a diene in a Diels–Alder [4 + 2] cycloaddition, leading to the formation of **53** (Scheme 13). This reaction is particularly interesting as it

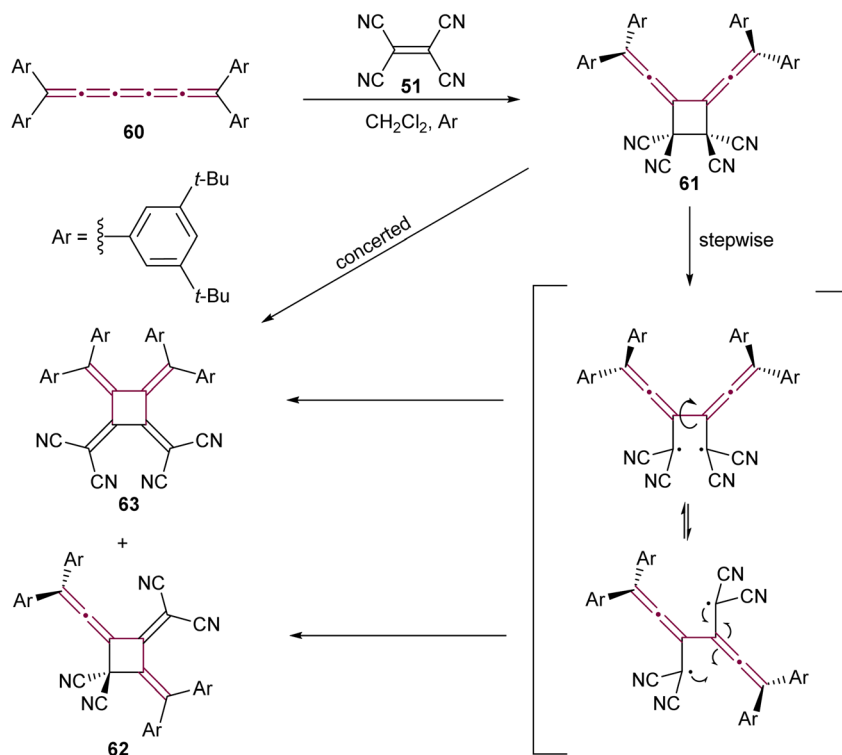
indicates that cycloadditions can involve the terminal α -bond of a [3]cumulene when sterically unencumbered. Supporting this reactivity pattern, a systematic theoretical study by Baroudi *et al.* reveals that both kinetic and thermodynamic factors govern the regioselectivity of cycloadditions in [3]cumulenes.⁹⁸

The Diederich group has demonstrated that a polarized push–pull [3]cumulene **18** exhibits a proacetylenic reactivity engaging in a CA–RE cascade with TCNE to give the stable zwitterion **54** (Scheme 14a).⁵⁶ This reactivity was later extended to a series of push–pull [3]cumulenes by the same research group.²⁹ In these studies, the proacetylenic nature of cumulenes was systematically evaluated through X-ray structural characterization and the measurement of rotational barriers about the cumulene axis. Remarkably, the rotational barriers were as low as 12 kcal mol⁻¹. These barriers were closer to those associated with sterically hindered single bonds rather than those typical of double bonds, which are expected to be significantly higher at around 65 kcal mol⁻¹ for ethylene.⁹⁹ Such results suggest that



Scheme 14 Reaction of TCNE with [3]cumulenes (a) **18**, (b) **55**, and (c) **55** under different conditions to form tetracenes **59**.^{100–103}



Scheme 15 Reaction of [5]cumulene **60** with TCNE.¹⁰⁶

highly polarized [3]cumulenes bear a greater structural resemblance to the single–triple–single bonding of acetylenes rather than the three contiguous double bonds that are characteristic of cumulenes.

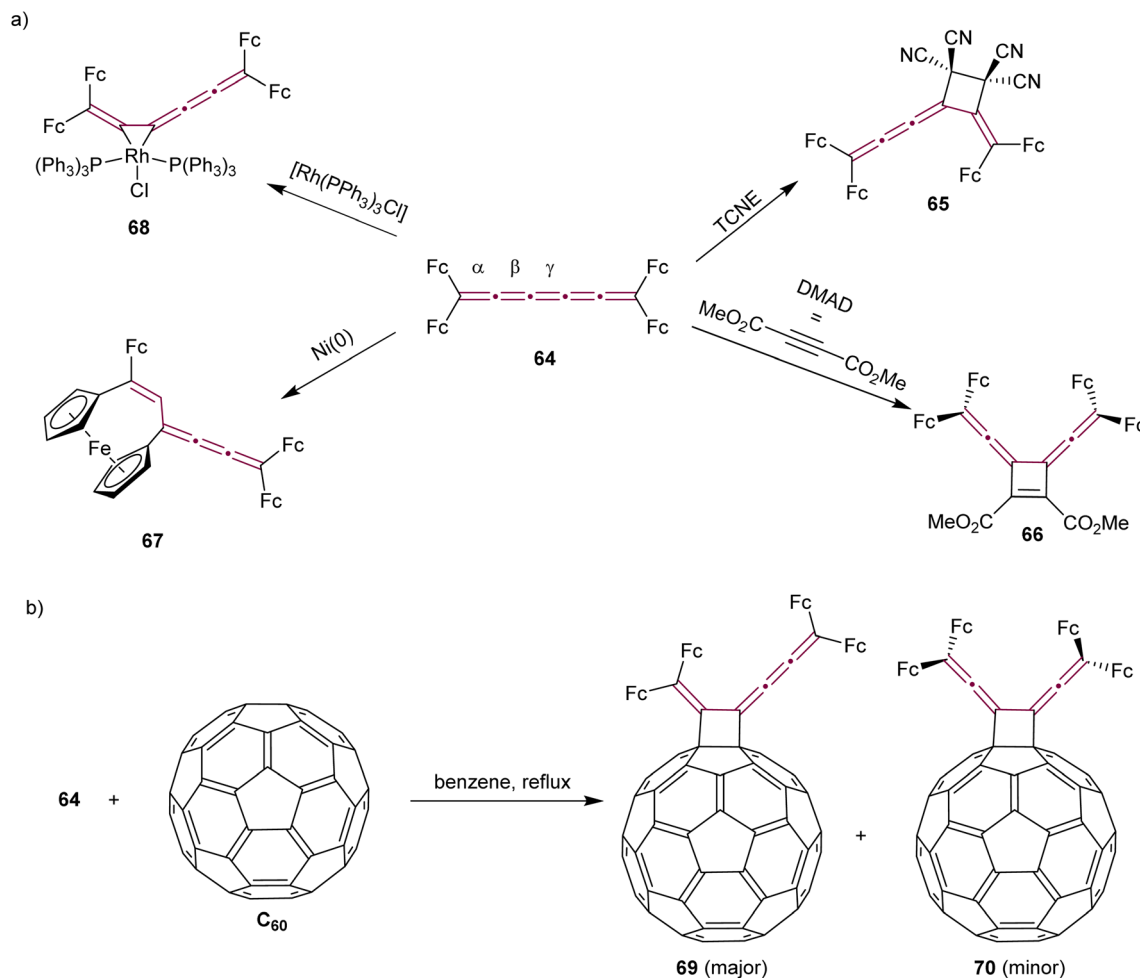
The proacetylenic reactivity of cumulenes with TCNE has been extended to symmetric, non-polar derivatives. Kawamura and coworkers showed that tetraaryl[3]cumulenes **55** undergo a facile reaction with TCNE in CH_2Cl_2 at ambient temperature to give cycloadducts **56** (Scheme 14b), which are relatively stable in nonpolar and aprotic solvents.¹⁰⁰ However, in polar solvents such as MeOH or MeCN, compounds **56** undergo an ionic transformation, leading to hydronaphthalenes **57** through ring opening followed by intramolecular cyclization of one of the aromatic substituents (Scheme 14b).¹⁰¹ Independently, Gawel *et al.* have detailed that the reaction of TCNE with tetraaryl[3]cumulenes **55** does not stop at **57** when performed in acetonitrile at elevated temperatures, but gives instead dihydrotetracenes **58** as the products of a multistep domino reaction (Scheme 14c).¹⁰² Detailed mechanistic investigations reveal that, after an initial CA–RE reaction between TCNE and the central β -bond of the [3]cumulene, two sequential 6-electron electrocyclizations furnish the tetracyclic carbon scaffold, which is followed by a 1,5-hydrogen shift and H_2 elimination to give **58**. In the presence of metallic copper, decyanation and aromatization are achieved by heating **58** neat, leading to cyanotetracenes **59**. Subsequently, a solution-based method has been developed using $\text{Cu}(\text{I})$ - N,N' -dimethylethylenediamine complex as a catalyst.¹⁰³ The resulting dicyanotetracenes **59** are strong fluorophores and undergo efficient singlet fission in the solid state.^{104,105}

Expanding on the reactivity of [n]cumulenes with TCNE, Januszewski *et al.* reported a complex outcome from the reaction between TCNE **51** and tetraaryl[5]cumulene **60** (Scheme 15).¹⁰⁶ The initial cycloaddition product, a vinylidene cyclobutane **61**, is isolable only at low temperatures. At room temperature, it undergoes a sequence of ring opening and closure, to form the cyclic [3]dendralene **62** and the electron-deficient [4]radialene **63**.

Tetraferrocenyl[5]cumulene **64**, studied by Bildstein and coworkers, reacts with TCNE *via* a formal [2 + 2] cycloaddition at the β -bond of the [5]cumulene to form the stable cyclobutane **65** (Scheme 16a).¹⁰⁷ The reaction of **64** with the electron-poor alkyne DMAD, however, takes place at the central γ -bond to produce cycloadduct **66** with a characteristic dark blue color. Attempts at [2 + 2] cyclodimerization with $\text{Ni}(0)$ to produce octaferrocenyl[4]radialenes lead to an unexpected outcome and the formation of a unique [3]ferrocenophane **67**. In the realm of coordination chemistry, **64** forms a stable complex with $[\text{Rh}(\text{PPh}_3)_3\text{Cl}]$. The [2 + 2] cycloaddition of cumulene **64** with C_{60} yields regioisomers **69** and **70**. The isomer **70** with D_{2h} symmetry, is the minor product that could not be isolated, whereas the green, air-stable isomer **69** is the major product, isolated in 68% yield (Scheme 16b).

In 1979, Bos and coworkers reported the heteroatomic photochemical [2 + 2] cycloaddition reaction between thioxanthenethione **71** and [3]cumulene **72**, which leads to the formation of cyclic thioether **73** as the major product (Scheme 17).¹⁰⁸ This report is a natural expansion of their earlier research on photochemical additions of thiocarbonyl compounds to acetylenes, allenes, and ketenimines *via* [2 + 2] and [4 + 2]





Scheme 16 Reactivity of tetraferrocenyl[5]cumulene **64** (Fc = ferrocenyl) (a) with metals and in cycloaddition reactions and (b) in the cycloaddition reaction with C_{60} .¹⁰⁷

cycloadditions.¹⁰⁹ In the reaction pathway, thioxanthenethione **71** in its triplet state preferentially attacks cumulene **72** at the terminal sp^2 -carbon, resulting in a stabilized biradical **74**. Subsequent ring closure produces the thietane derivative **73**. Interestingly, the formation of a minor isomeric byproduct, cyclobutanethione derivative **75**, can be explained by the thioxanthenethione **71** attack at C2 of cumulene **72**, followed by ring closure of the biradical **76** to give thione **77**. The thione **77** then undergoes a rearrangement, resulting in the formation of the final cyclobutanethione **75** (Scheme 17).

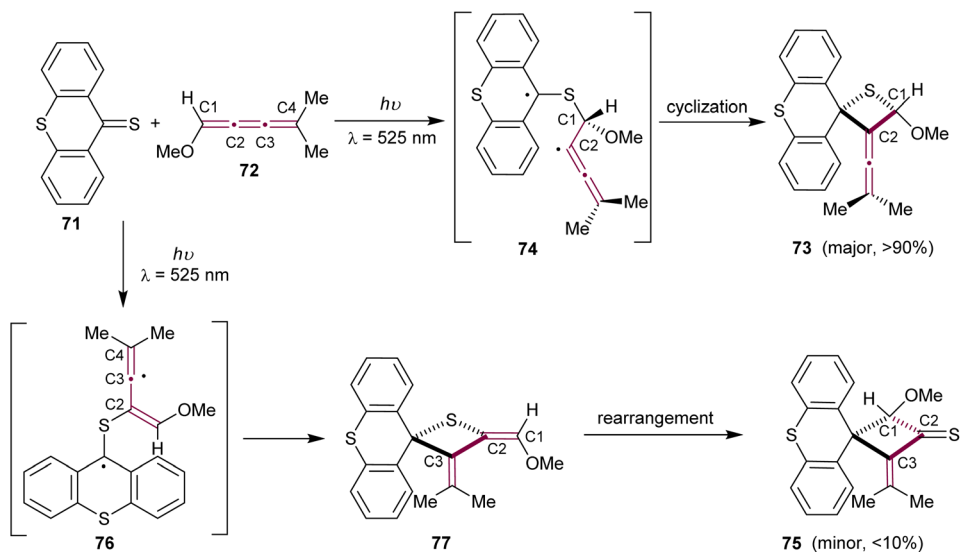
In the context of research dedicated to the mechanism underlying the antitumor efficacy of the endiye antibiotic neocarzinostatin, a series of studies have examined the intramolecular cyclization of enyne[3]cumulenes.^{110–112} The Hiram group reported on the intramolecular electrocyclic reactions of enyne[3]cumulene **78**.¹¹³ Under thermal conditions (in deoxygenated 1,4-cyclohexadiene at 80 °C), **78** gave a mixture of styrene **79** (19%) and benzocyclobutane **80** (21%) as products of Bergman-type cyclization and [2 + 2] cycloaddition pathways, respectively (Scheme 18a). With a meticulous mechanistic investigation, this study shows that the reactions indeed follow

first order kinetics that are characteristic of intramolecular reactions and that radical species are involved.

A few years later, Rodríguez and coworkers have explored the intramolecular cycloadditions in enyne[3]cumulenes **81** and **82**.¹¹⁴ In these cases, the reaction pathway depends on the terminal substitution of the alkyne (Schemes 18b and c). [3]Cumulene **81**, bearing a terminal alkyne, converts efficiently into naphthalene **83** via a Bergman-type cyclization, whereas a trimethylsilyl-protected alkyne **82** gives cyclobutene **84** as a result of an intramolecular [2 + 2] cycloaddition reaction (Schemes 18b and c). Both computational investigations and experimental observations suggest that steric effects play a crucial role in the intramolecular cyclization of enyne [3]cumulenes, with bulky substituents promoting [2 + 2] cycloaddition over Bergman-type cyclization.

The potential of cumulenes as synthons was further demonstrated by Guan and Shi, who reported the phosphine-mediated [3 + 2] cycloaddition reactions of [3]cumulene **85** (Scheme 19).¹¹⁵ In the presence of PBU_3 , [3]cumulene **85** reacts with substituted methylidenemalononitrile **86** or *N*-tosylimine **87** to give substituted cyclopentene **88** and pyrrolidine **89**,





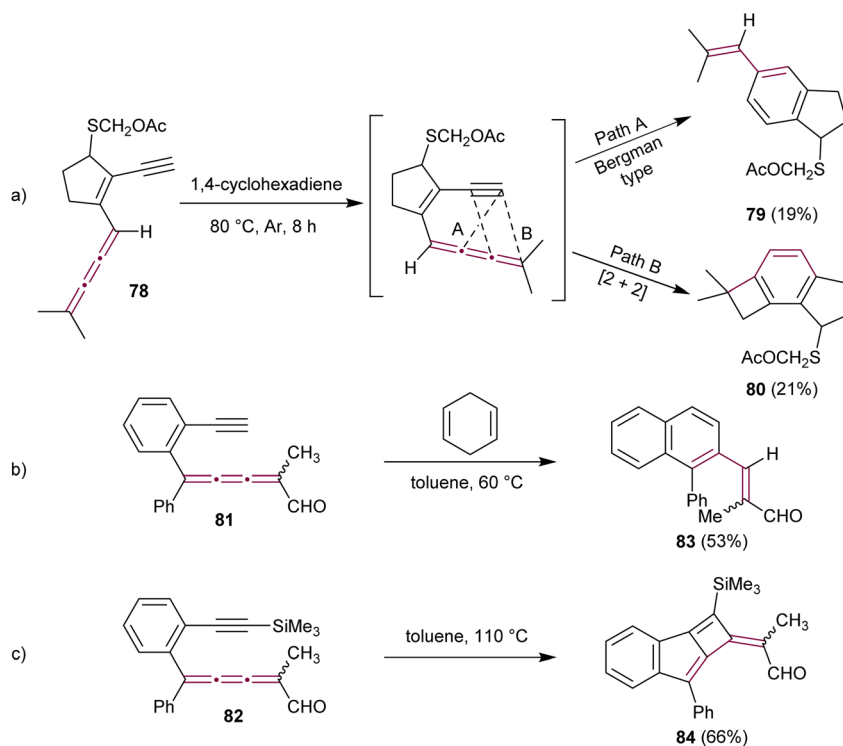
Scheme 17 Photochemical [2 + 2] cycloaddition of thioxanthenethione **71** to [3]cumulene **72**.¹⁰⁸

respectively. A plausible reaction mechanism was proposed based on previously reported phosphine-catalyzed [3 + 2] cycloadditions.¹¹⁶ The first step involves the nucleophilic attack of PBU_3 on **85** to produce a zwitterionic intermediate that attacks the electrophile **86** (or **87**), followed by an intramolecular conjugate addition and elimination of phosphine.

Another report showed that the tributylphosphine-catalyzed [3 + 2] cycloaddition of **85** with the aromatic aldehyde **90** leads to the formation of substituted furan derivative **91**.¹¹⁷

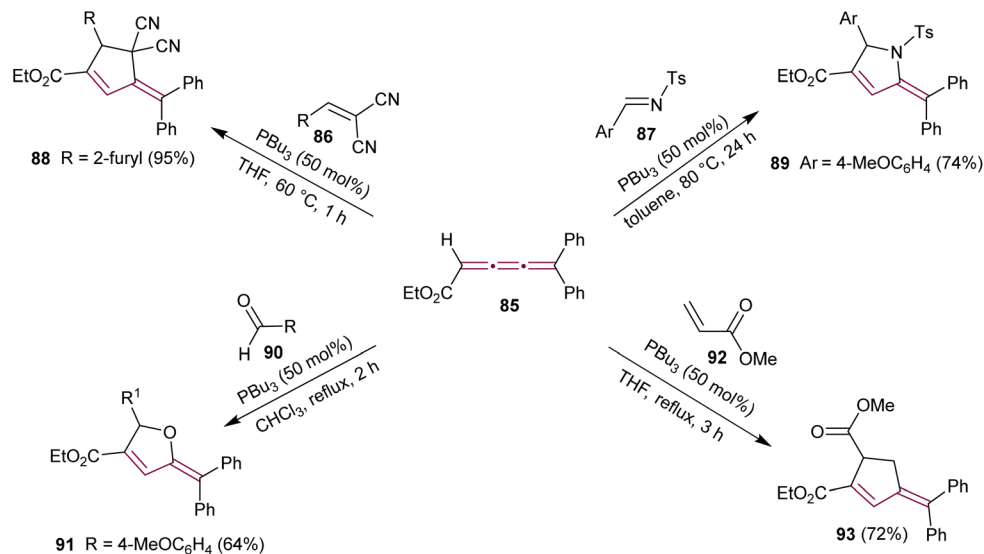
Additionally, the reaction of **85** with the α,β -unsaturated carbonyl esters, such as **92**, yields the substituted cyclopentene **93** (Scheme 19).

In an effort to explore the reactivity of 1,3-diaza-2-azoniaallene salt **94**, Jochims and coworkers examined the reaction with [3]cumulene **95**, leading to the formation of allene **96** (Scheme 20).¹¹⁸ The cycloaddition occurs exclusively at the α -bond of the [3]cumulene moiety. Computational studies indicate that **96** is 3–11 kcal mol⁻¹ more stable than any alternative



Scheme 18 Intramolecular cyclization reactions of enyne[3]cumulenes: (a) **78** via a Bergman-type cyclization (Path A) or [2 + 2] cycloaddition (Path B), (b) Bergman-type cyclization reaction of **81**, and (c) [2 + 2] cycloaddition reaction of **82**.^{113,114}





Scheme 19 Phosphine-mediated [3 + 2] cycloaddition reactions of ethyl 5,5-diarylpenta-2,3,4-trienoates **85**.^{115,117}

regioisomer that could be formed upon addition to the central β -bond of **95**. Interestingly, upon recrystallization, **96** transforms into the triazolium salt **97** with the loss of isobutene. According to the proposed mechanism, this transformation proceeds *via* elimination of Me_3C^+ from the 4-position of **96**, followed by proton transfer from Me_3C^+ to the intermediate structure, ultimately yielding isoprene and **97**.

Organometallic reactions

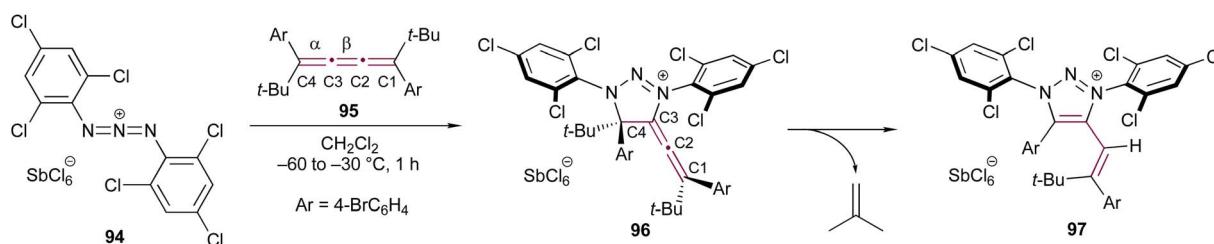
Among the many tools available in modern organic synthesis, organometallic reactions stand out for their versatility in constructing complex molecular architectures.¹¹⁹ Within this context, cumulenes are intriguing, yet underexplored, substrates, and their distinct bonding pattern and electronic structure impart unique reactivity patterns that hold promise for organometallic transformations. Due to their accessible π -systems and their suitable energy levels, cumulenes react readily with many metals, ranging from alkali to late transition metals. Among many unsaturated hydrocarbons, [n]cumulenes have been widely used as ligands in coordination chemistry;^{53,120–141} however, here, we focus mainly on the utility of cumulenes as synthons in organic synthesis. Hence, in this section, we explore the nuanced reactivity of cumulenes mediated by metal centers.

The work of Zweig and Hoffmann marked the first reported metal-induced reaction involving a cumulene, detailing the

reduction of tetraphenyl[3]cumulene **1** using metallic sodium or potassium in an ethereal solution (Scheme 21).¹⁴² This reaction initially forms a radical anion intermediate, which subsequently undergoes reduction to yield dianion **98**. The dianion appears as a transient brown suspension that gradually fades over time. Electron spin resonance (ESR) results reveal a weak signal corresponding to the radical intermediate, which diminishes and disappears within two to three hours. The disappearance of the ESR signal suggests conversion to ESR-silent dianion **98** or other closed-shell species. Treatment of dianion **98** with alkyl halides leads to functionalization at propargylic positions affording **99** (Scheme 21). This study was followed by the study of Day and coworkers who explored the scope of alkali metal-induced reactions of **1** and other tetraaryl[3]cumulenes.^{143,144}

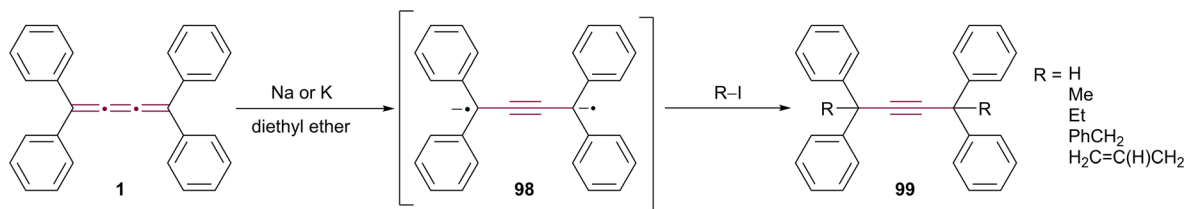
Dianions of tetraphenyl[3]cumulene **1** have been extensively studied by Kornacki and Kemula using polarographic methods as well as structurally characterized by X-ray crystallography by Bock and coworkers.^{145–148} Recently, in their collaborative efforts, the groups of Tykwinski and Petrukhina isolated and characterized complexes of [3]- and [4]cumulenes with alkali metals, showing structural changes in cumulenic cores upon reduction.^{149,150}

Lithiation of the alkyl-substituted tetramethyl[3]cumulene **12** shows different selectivity than tetraphenyl derivative **1**. Ando and coworkers postulated that treating **12** with excess Li metal in Et_2O produces organometallic intermediate **100**



Scheme 20 Reaction of [3]cumulene **95** with 1,3-diaza-2-azoniaallene salt **94**.¹¹⁸





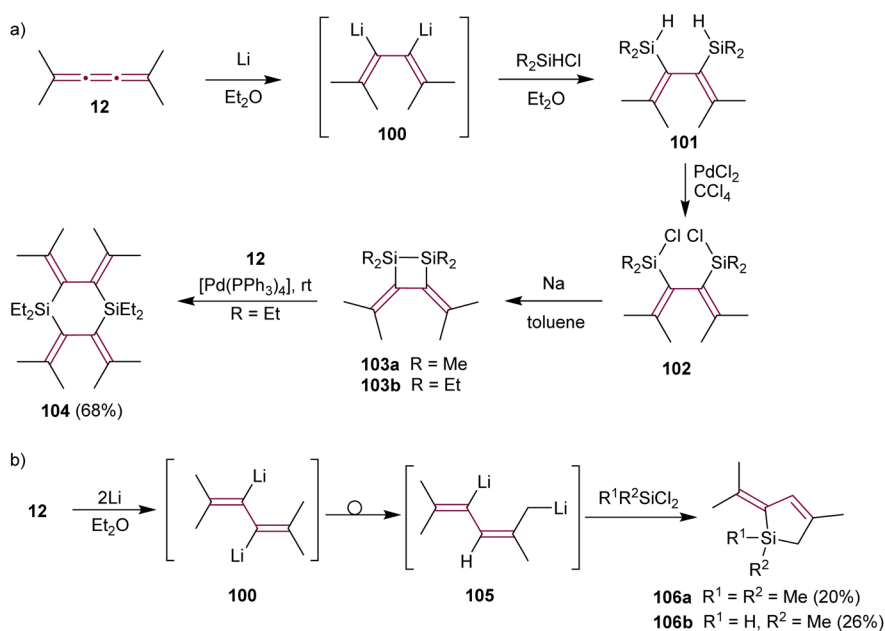
Scheme 21 Formation and reactivity of anionic species derived from tetraphenyl[3]cumulene **1**.¹⁴²

(Scheme 22a). Trapping this dianion with a chlorosilane (R_2SiHCl) affords the bis-silylated diene **101** via Li/Si exchange. The subsequent reaction of **101** with $PdCl_2$ yields chlorosilanes **102**. Dienes **102**, upon reduction with sodium, give disilacyclobutanes **103a** or **103b**.¹⁵¹ Such disilanes are useful reagents in transition metal-catalyzed double silylation of C–C multiple bonds.¹⁵² Interestingly, the nature of alkyl substituents of the silane affects the yields of doubly silylated products and subsequent disilane metathesis. While methylsilane derivative **103a** displays high reactivity and requires storage in dilute solution to prevent polymerization, ethyl-functionalized **103b** exhibits remarkable stability for several months even in neat form and at room temperature. Ethylsilane **103b** reacts with [3]cumulene **12** in a Pd(0)-catalyzed reaction producing cyclic silane **104** (Scheme 22a). Subsequently, Maercker and coworkers report that intermediate **100** rearranges to the 2,5-dilithiated isomer **105**, which is then transformed to dihydrosilols **106** (Scheme 22b).¹⁵³

In 1990, Ziegler have detailed a study on the Cu-mediated isomerization of bromo[3]cumulene **107** to bromoenyne **108**, followed by vinyl radical cyclization that leads to the synthesis of substituted cyclopentene **109** (Scheme 23).¹⁵⁴ This enyne moiety is a key component of the strained fused ring system found in the antitumor antibiotic neocarzinostatin.^{110–112} Interestingly, treatment of **107** with AIBN and $n-Bu_3SnH$ leads

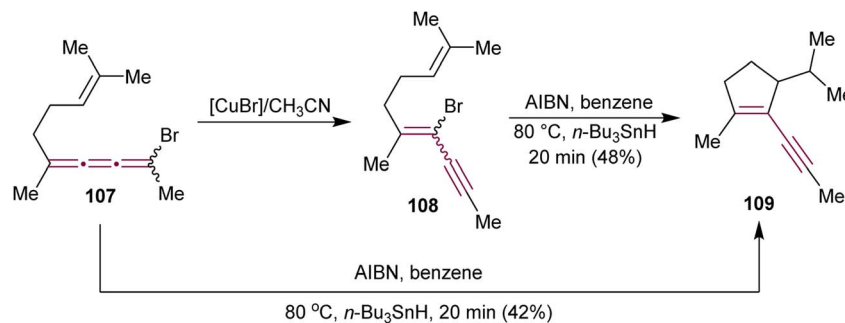
directly to **109** via 5-(π -endo)-*exo*-trig cyclization. Both **107** and **108** convert to **109** via a common planar radical intermediate.

The pioneering Pd-catalyzed arylation of cumulenes was demonstrated by Jones and coworkers. They demonstrated that [3]cumulene **1** reacts with iodobenzene through a Pd-mediated catalytic cycle to give substituted benzofulvene **110** in 97% yield (Scheme 24a).¹⁵⁵ This transformation is proposed to commence with a carbopalladation across the β -bond of the cumulene moiety and then proceed through a cyclopalladation/C–H activation step. Reductive elimination to **110** concludes the final step. Kan and coworkers applied a similar Pd-catalyzed reaction to convert [3]cumulene **111** into benzofulvene **112** (Scheme 24b). The scope of this reaction is broadened to provide good yields using a range of aryl halides containing either electron-donating or -withdrawing groups.^{156,157} Remarkably, the conjugated carbonyl group in **111** does not alter the course of the reaction: the aryl-palladium intermediate still inserts regioselectively into the β -double bond. Subsequent C–H activation with a neighboring phenyl group leads to the formation of 2,3-benzofulvene derivatives **112** with a substituted exocyclic alkene. This domino process deviates from allene formation initially anticipated by the authors. Aryl iodides containing both electron-donating and -withdrawing groups produce corresponding products in good yields.



Scheme 22 Lithium-induced reactions of tetramethyl[3]cumulene **12** with (a) chlorosilanes and (b) dichlorosilanes.^{151,153}

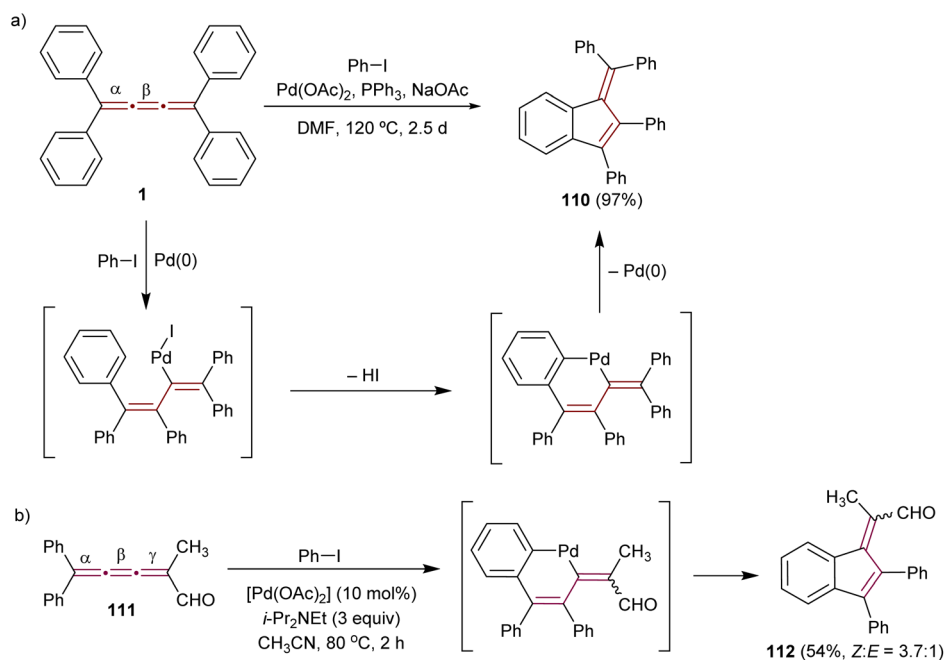


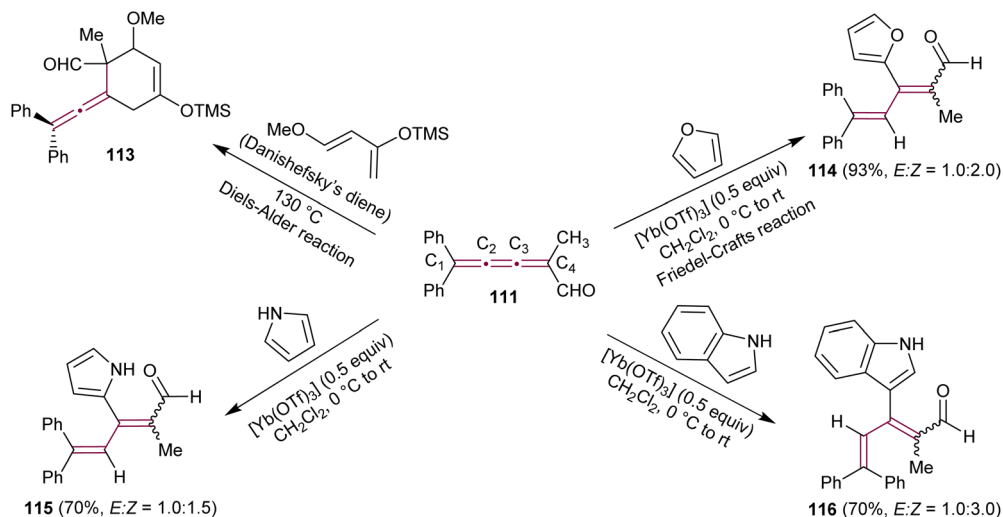
Scheme 23 Cyclization of bromo[3]cumulene **107** and bromoene **108** to cyclopentene **109**.¹⁵⁴

Kan's group expanded the use of [3]cumulene **111** as a synthon (Scheme 25).¹⁵⁸ In a Lewis-acid catalyzed Diels–Alder reaction, cumulene **111** reacts efficiently with Danishefsky's diene producing an unstable allenyl cycloadduct **113**. The reaction of **111** with cyclopentadiene is also described in detail (not shown). Notably, Diels–Alder reactions occur selectively at the C3–C4 double bond. Further investigations revealed the successful Friedel–Crafts reactions with furan, catalyzed by the Lewis acid $\text{Yb}(\text{OTf})_3$, yielding tetrasubstituted, conjugated diene **114** in a 93% yield. This reaction was proposed to proceed through the conjugate addition at C3 of the cumulene, followed by subsequent protonation at C2. The Friedel–Crafts reactions with nitrogen-containing heteroaromatics, such as pyrrole and indole, were also reported. In all cases, the reactions proceed smoothly, yielding the corresponding dienes **115** and **116** in excellent yields, with no evidence of the Diels–Alder product or the corresponding allene derivative. To validate the reactivity of **111**, the authors also tested the reaction with other nucleophiles, such as thiolates. Similar to the Friedel–Crafts reaction, the conjugate addition of the thiolate anion occurs at C3,

followed by protonation at C2, resulting in the conjugated diene with same regioselectivity as **114** and **115**.¹⁵⁸ The conjugate addition–protonation sequence can be considered as a formal 3,4-addition, forming tetrasubstituted conjugated dienes. This reactivity stands in stark contrast to the Diels–Alder reaction, which occurs at the C3–C4 bond of **111**, leading to tetrasubstituted allene derivatives, such as **113**. As expected, this reactivity is influenced by electronic and steric factors inherent for each double bond. The electron-withdrawing effect of the formyl group and the steric hindrance of the diphenyl moiety are postulated as key factors contributing to this unique reactivity.

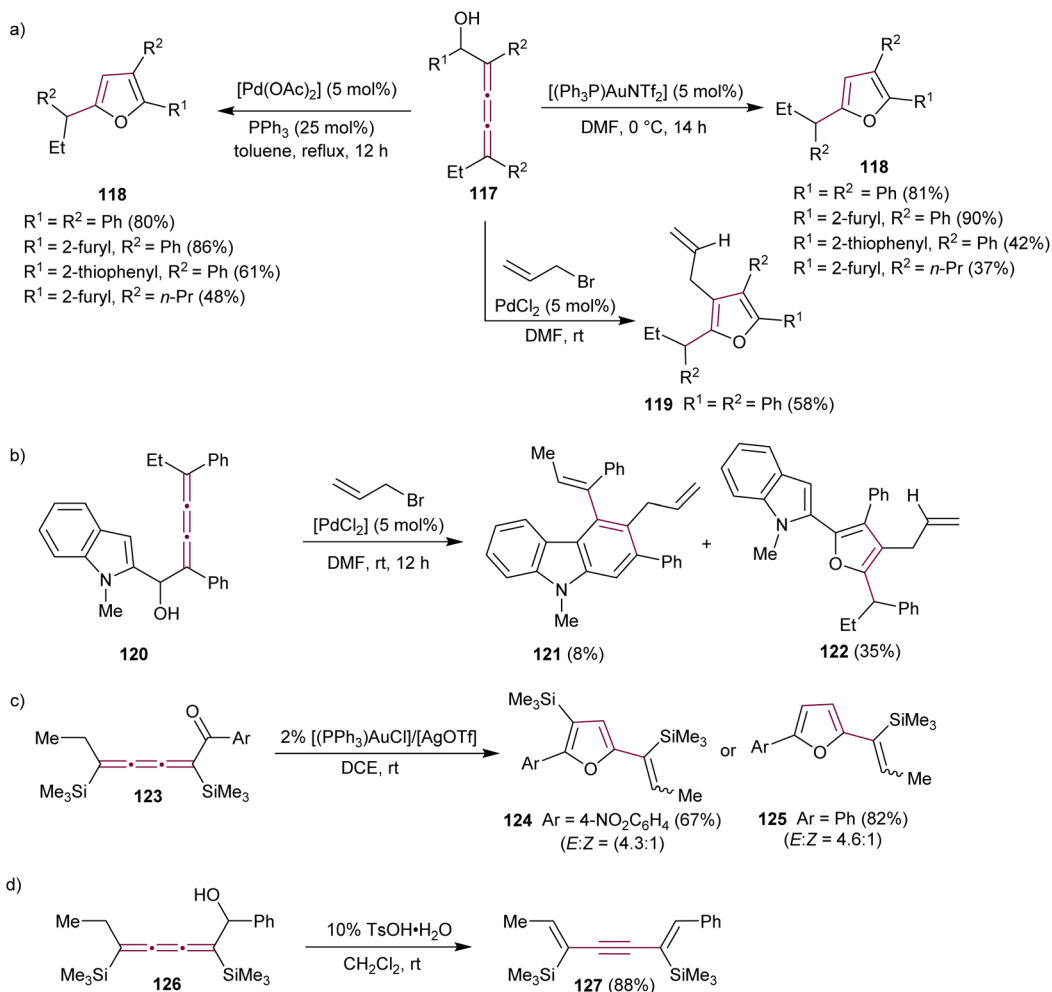
Alcaide *et al.* reported the effect of several metal catalysts on the cycloetherification reaction of 2,3,4-trien-1-ols **117**, motivated by the prevalence of oxacyclic structures (*e.g.*, furan derivatives) in many biologically active natural compounds (Scheme 26a).^{159,160} The study identified several distinct reaction conditions that can be implemented to make tri- and tetrasubstituted furans **118** and **119** in a controlled manner.¹⁶⁰ A carbocyclization/coupling sequence was also explored in indole-

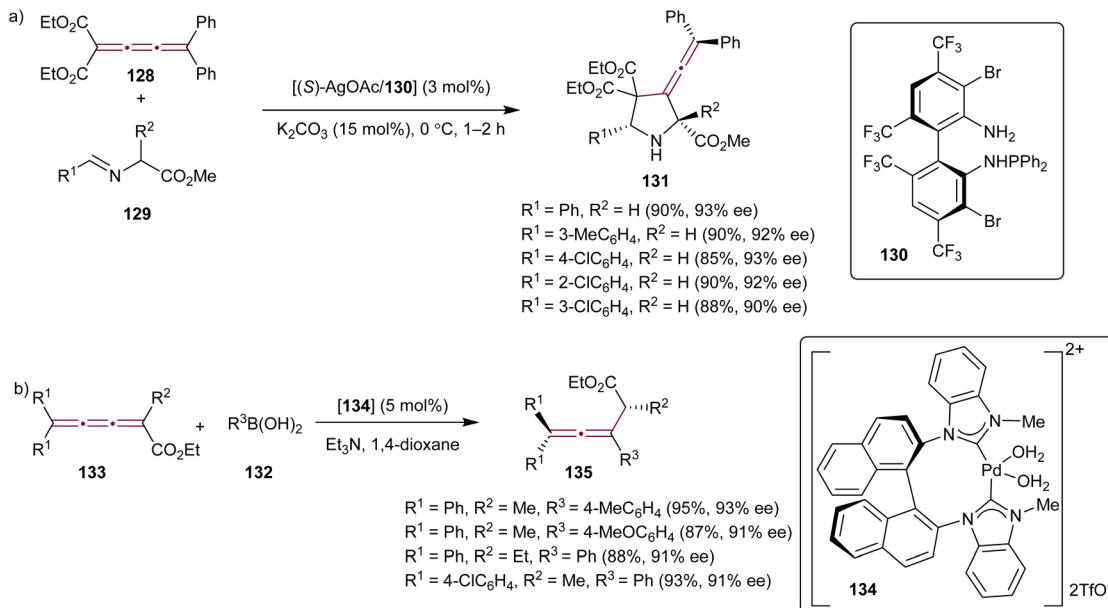
Scheme 24 Carbopalladation reactions of [3]cumulenes (a) **1** and (b) **111**.^{155–157}

Scheme 25 Reactions of cumulenal **111** in Diels–Alder and Friedel–Crafts reactions.¹⁵⁸

substituted [3]cumulene **120**, leading to the formation of 2,3,4-trisubstituted carbazoles **121** as the minor product together with furan **122** (Scheme 26b).

Liu and coworkers showed that cycloisomerization of [3]cumulenone **123** catalyzed by gold(i) complexes yields tri- and disubstituted vinyl furans **124** and **125** as the major and minor

Scheme 26 Reactions of (a) [3]cumulenol **117**, (b) [3]cumulenol **120**, (c) [3]cumulenone **123**, and (d) [3]cumulenol **126**.^{160,161}

Scheme 27 Stereoselective organometallic reactions of [3]cumulenes (a) **128** and (b) **133**.^{162,164}

products, respectively (Scheme 26c).¹⁶¹ The same study reported that Brønsted acid-catalyzed isomerization of [3]cumulenol **126** to 1,5-dien-3-yne **127** is a mild and selective synthetic method for entry into this useful class of compounds (Scheme 26d).

In the context of constantly growing demand for synthetic methods capable of producing enantiopure pharmaceutical compounds, the exploration of organometallic reactions involving cumulenes has also been targeted with the objective of achieving high stereoselectivity in the construction of complex molecular systems. In 2011, Xue *et al.* described a catalytic asymmetric 1,3-dipolar cycloaddition reaction between tetrasubstituted cumulene **128** and azomethine ylides **129** (Scheme 27a).¹⁶² This transformation was catalyzed by the Ag(I) complex with the ligand TF-BiphamPhos **130** and gave 3-vinylidene-pyrrolidine derivatives **131** in high yields, with enantioselectivities reaching 93% ee.

Catalytic, asymmetric conjugate addition of organoboron reagents to alkenes or allenes has been a well-established and efficient approach for creating chiral, enantioenriched molecules from simple, often achiral, building blocks.¹⁶³ Shi and coworkers have expanded this methodology to include cumulenes. The authors reported a conjugate addition of arylboronic acids **132** to cumulenes **133** (Scheme 27b) catalyzed by the axially chiral *N*-heterocyclic carbene Pd(II) complex **134**.¹⁶⁴ This transformation resulted in the formation of allenes **135** with good to excellent yields and moderate to good enantioselectivities.

Conjugated bis[3]cumulene **136** is composed of two cumulene units linked by a single bond (Scheme 28a). Owing to the presence of numerous unsaturated carbon atoms within their structure, they are excellent candidates for complex domino reactions. Konishi *et al.* reported the Lewis-acid catalyzed cycloisomerization of conjugated bis[3]cumulene **136** that leads to the formation of functionalized pentalenes **137** (Scheme 28a).¹⁶⁵ The rigid acenaphthene backbone enables effective communication between

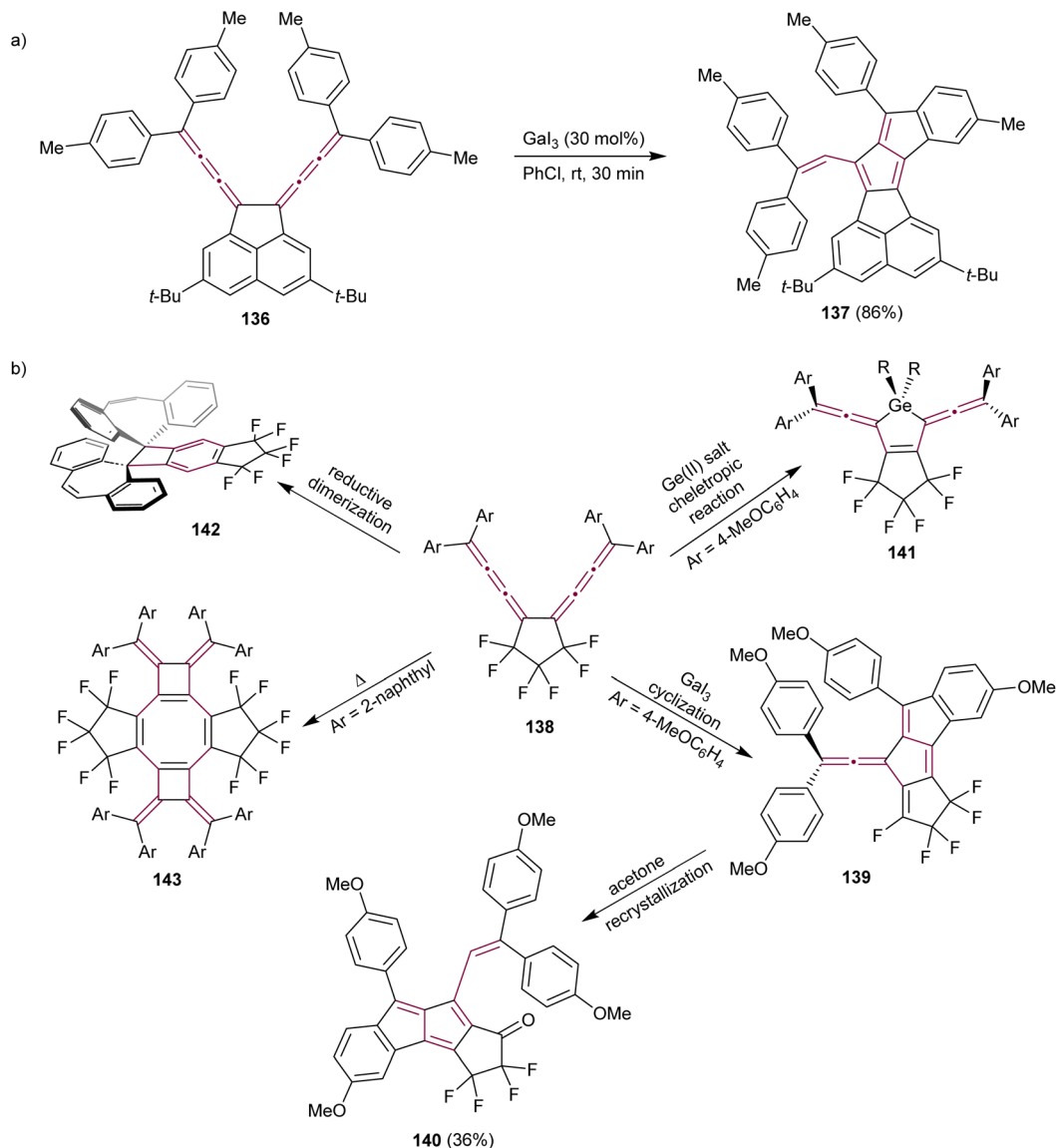
the two exocyclic [3]cumulenes, facilitating the one-pot construction of pentalenes **137** through the catalytic cycle involving Lewis acid, such as GaI₃. The resulting diareno[a,f]pentalene derivative **137** is a stable purple solid with a nearly planar structure, as shown by X-ray crystallographic analysis.

The same research group have investigated the reactivity of hexafluorocyclopentane-bridged bis[3]cumulenes **138** (Scheme 28b).¹⁶⁶ Electron-withdrawing fluorine atoms enhance the stability of cumulenes and the flexible aliphatic cyclopentane minimizes steric hindrance relative to the acenaphthene backbone, resulting in new reactivity patterns. The reaction of **138** with the Lewis acid GaI₃ results in extended fulvene **139**. Interestingly, fulvene **139** transforms to pentalene **140** during crystallization in acetone when functionalized with electron-donating aryl substituents, such as *p*-anisyl. Biscumulenes **138** serve as ligands when treated with GeCl₂ and form stable germacycles **141**. When the cumulene bears dibenzocycloheptatriene as an endgroup, a reductive Bergman-type cyclization takes place to produce benzocyclobutane **142**, featuring an extraordinarily long single C–C bond of 1.75 Å between dibenzocycloheptatriene moieties. They also showed that the bis[3]cumulene **138** endcapped with 2-naphthyl groups undergoes thermal dimerization to form the functionalized cyclooctatetraene derivative **143**. This study beautifully demonstrated the potential of cumulenes as synthons, showing that from one cumulene-based scaffold, a range of complex π -conjugated structures can be synthesized by simply altering the reagents and reaction conditions.

Miscellaneous reactions

In the previous sections, we provide detailed discussions on cyclooligomerization, cycloaddition, and organometallic reactions involving odd [*n*]cumulenes. This section aims to integrate and examine the diverse array of additional reaction types involving cumulenes that do not fit into the other groups. They





Scheme 28 Catalytic reactions of bis[3]cumulenes (a) **136** and (b) **138**.^{165,166}

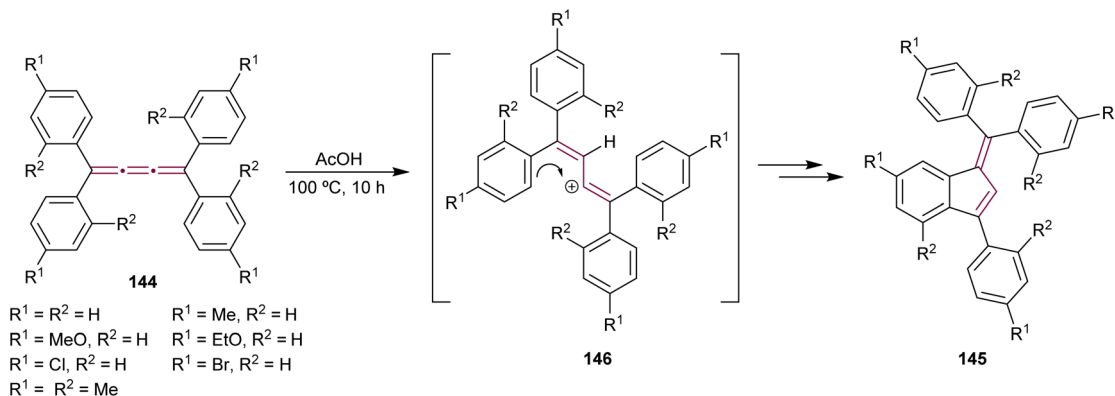
include reactions involving acids, bases, and halogens, as well as epoxidation reactions and various other processes. This will further highlight the synthetic utility and functional versatility of $[n]$ cumulenes in organic chemistry.

Benzofulvenes are a valuable class of compounds widely used as precursors in total synthesis, organometallic chemistry, and materials science.^{167–169} Given their synthetic utility, numerous strategies have emerged for assembling the benzofulvene scaffold. As discussed earlier, synthesis of benzofulvenes can be achieved through the Pd-catalyzed arylation of [3]cumulenes (Scheme 24).^{155–157} However, already in 1950, Brand reported that tetraaryl[3]cumulenes **144** cyclize to benzo[*d*]fulvene **145** upon heating in acetic acid at 100 °C (Scheme 29).¹⁷⁰ The reaction likely proceeds *via* protonation of cumulene, producing the carbocation intermediate **146**, which then undergoes an intramolecular Friedel–Crafts arylation to form a five-membered ring. The structure of **145** and the reaction mechanism were supported by subsequent studies, which

demonstrated that treatment of various [3]cumulenes with acid consistently produced similar structures.^{171,172}

Diederich and coworkers have reported the synthesis of cyanated benzo[*c*]fluorenes **147** from the reaction of tetraaryl [3]cumulenes **55** with 2,3-dichloro-5,6-dicyano-1,4-benzoquinone (DDQ), a common oxidizing agent (Scheme 30a).¹⁷³ The proposed mechanism begins with a single electron transfer from the cumulene to DDQ, followed by intramolecular cyclization to form the fulvene core. Cycloaddition of this intermediate with DDQ results in **148**, which was isolated and its structure confirmed by single-crystal X-ray diffraction analysis. Subsequent elimination of 2,3-dichloromaleate yields the final product, fulvenes **147**. When substituted with electron-donating anisyl groups, cumulene **149** undergoes the same cyclization step to a benzofulvene intermediate, but the reaction is terminated by the addition of DDQ to produce benzofulvene **150** (Scheme 30b).



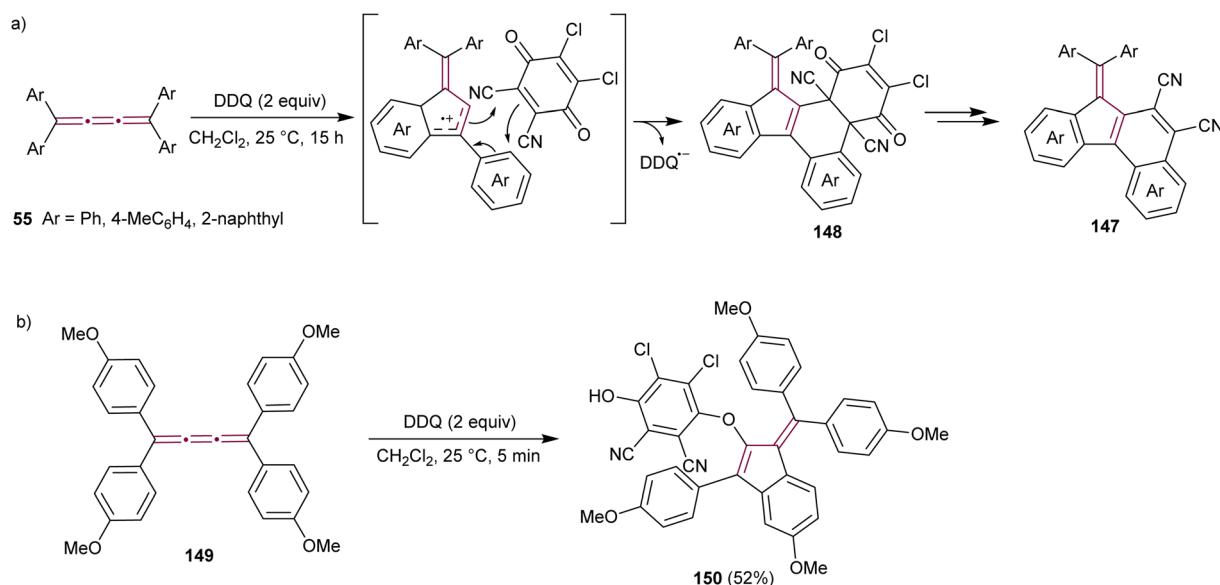
Scheme 29 Acid-mediated rearrangement of tetraaryl[3]cumulenes **144** to benzo[d]fulvenes **145**.¹⁷⁰

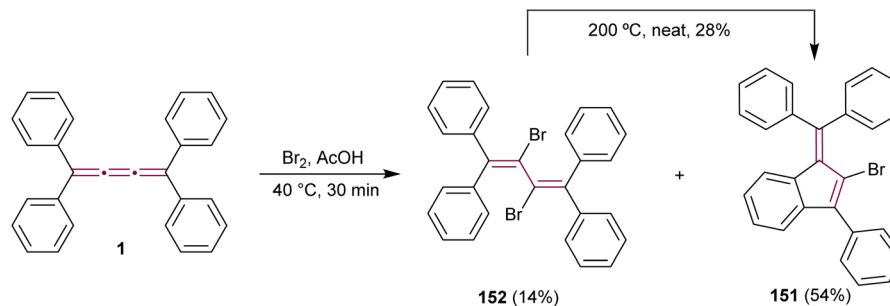
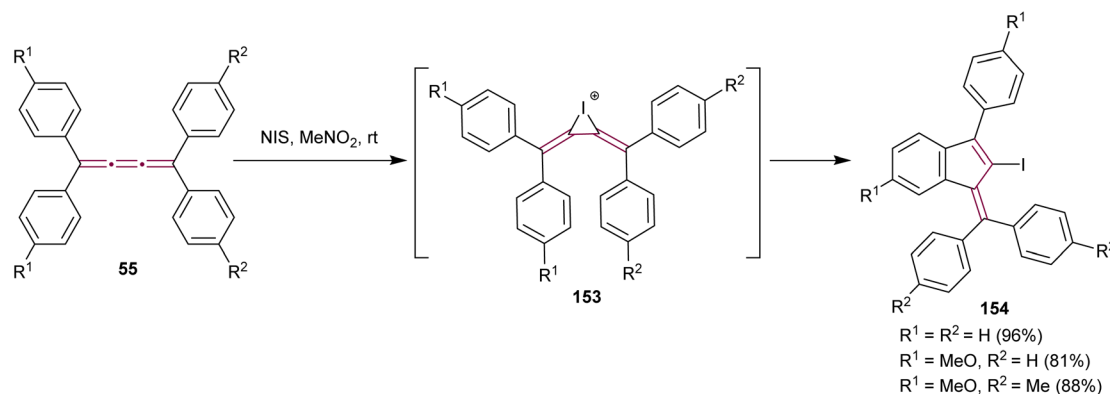
The outcome of halogenation reactions of cumulenes differs substantially from that of alkenes. For instance, bromination of tetraphenyl[3]cumulene **1** yields a mixture of products with compositions that vary depending on the reaction conditions.³⁸ Woliński reports that, in a polar protic solvent such as acetic acid, the primary product of reaction with bromine is benzo[fulvene] **151**, implicating polar or ionic intermediates in its formation. Conversely, carrying out the reaction in an apolar solvent like CCl_4 and catalyzed by benzoyl peroxide favors the formation of butadiene **152**, which represents the expected product upon bromination of a double bond.³⁸ Additionally, 2,3-dibromobutadiene **152** can be transformed into fulvene **151** upon heating to 200 °C in the absence of solvent (Scheme 31).

In 2021, Yagishita *et al.* described the synthesis of the benzo[fulvene] framework **154** *via* electrophilic iodocyclization of tetraaryl[3]cumulenes **55** (Scheme 32).¹⁷⁴ The reaction utilizes *N*-iodosuccinimide (NIS) to generate iodonium intermediates **153**, which undergo rearrangement forming iodobenzo[fulvenes] **154**. The cyclization step occurs regioselectivity on a more electron-rich aromatic ring when cumulenes are unsymmetrically

substituted. Under optimized conditions, the equivalent reaction with NBS instead of NIS gives the analogous bromo-derivatives of benzo[fulvenes] **154**. The NIS transformation is subsequently shown to proceed efficiently under solvent-free mechanochemical conditions.¹⁷⁵

The formation of highly functionalized fulvenes through the reaction of tetraaryl[5]cumulene **155** with elemental iodine is depicted in Scheme 33a. This reaction yields diiodofulvene **156** and iodofulvene **157** in varying ratios, depending on reaction conditions.¹⁷⁶ Mechanistic studies suggest that iodonium ion intermediate **155A** is generated by the reaction of **155** with iodine at the β -bond (Scheme 33b). The subsequent nucleophilic attack of the iodide ion on **155A** produces diiodide **155B**, which further reacts with another molecule of iodine to form the iodonium intermediate **155C**. Subsequent cyclization, migration of the aryl group, and elimination of iodonium give the final diiodofulvene **156**. When the reaction time is increased, in either $MeNO_2$ or benzene, the authors noted a decreased yield of **156** and increased yield of **157**. This suggests that compound **157** is generated from compound **156**,

Scheme 30 Reactions of DDQ with tetraaryl[3]cumulenes (a) **55** and (b) **149**.¹⁷³

Scheme 31 Bromination of tetraphenyl[3]cumulene **1**.³⁸Scheme 32 Synthesis of iodofulvenes **154** from tetraaryl[3]cumulenes **55**.¹⁷⁴

presumably by the reaction of **156** with HI generated *in situ*. The synthetic utility of iodofulvenes **154** and **157** is showcased in both reports through functionalization *via* a cross-coupling reactions that produced unique π -expanded fulvenes.¹⁷⁶

Vinylidene cyclopropanes are stable molecules, characterized by an exocyclic allene moiety linked to a cyclopropane ring and provide synthetic access to a range of unique hetero- and polycyclic compounds.¹⁷⁷ Shi and Li reported a triflic imide-catalyzed reaction cascade between diarylvinylidene cyclopropanes **158** and [3]cumulene **85** (Scheme 34a).¹⁷² The final product of this cascade annulation reaction varies depending on the substitution of the vinylidene cyclopropane substrate. Vinylidene cyclopropanes **158** substituted with electron-donating groups typically favor the formation of **159**, while those substituted with electron-withdrawing groups predominantly yield **160**. The proposed mechanism involves activation of [3]cumulene **85** by Brønsted acid $\text{ Tf}_2\text{NH}$, leading to the formation of a carbocation intermediate and subsequent nucleophilic attack by diarylvinylidene cyclopropane **158** (Scheme 34a). This is followed by an intramolecular cyclization, which yields the product **159** or **160** depending on the nature of aryl groups.

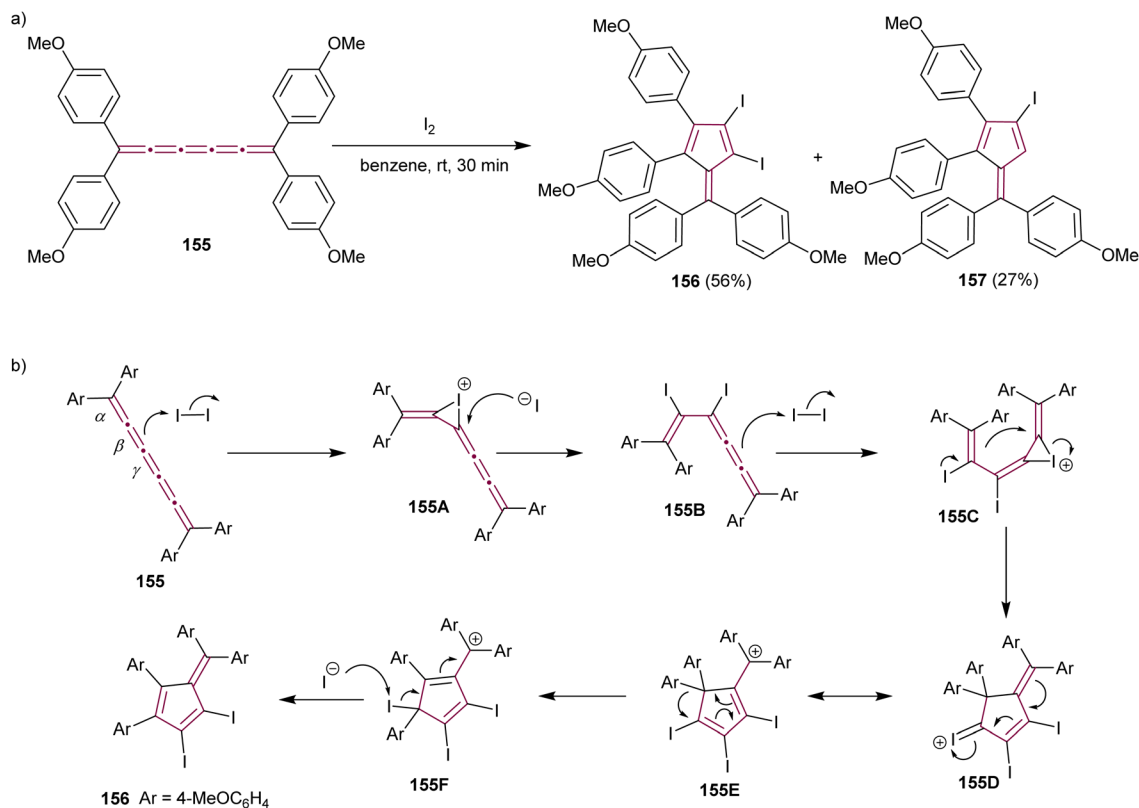
Interesting outcomes have been reported by Ando and coworkers from the reaction of tetraaryl[3]cumulenes **55** with elemental sulfur and selenium (Scheme 35).¹⁷⁸ After heating the [3]cumulenes **55** with elemental sulfur for several hours, macrocyclic products **161** are obtained. [3]Cumulenes **55** are inert to elemental selenium and required base 1,8-diazabicyclo[5.4.0]undec-7-ene (DBU) to give 1,2,5-triselenophane

derivatives **162**. When compounds **161** and **162** are heated for extended periods in the presence of DBU, the major product is a 1,3-butadiene in both cases, the product of reduction. However, more intriguing are the minor products, benzothio-phenone **163** and benzoselenophene **164**, which form *via* intramolecular cyclization cascades. When the [3]cumulene is substituted with adamantylidene groups, the product of the sulfurization reaction is analogous to **161**.¹⁷⁹

As later reported by Ando and coworkers, the sulfurization of unsymmetrical [3]cumulene **165** proceeds quite differently compared to symmetric tetraaryl- (**55**) or adamantylidene-substituted derivatives.^{178–180} When **165** is heated neat with molten sulfur or in a solution of diphenyl ether, the exclusive product of this reaction is 1,2-dithiolo-1,2-dithiole **166**. However, performing the sulfurization in DMF leads to 1,2-dithiole **167** as the major product (Scheme 36a). Interestingly, upon heating to 250 °C in benzene under N_2 , 1,2-dithiolo-1,2-dithiole **166** undergoes thermolysis producing an unusual polycyclic scaffold **168** (Scheme 36b). In contrast, upon UV irradiation (360 nm), **166** rearranges to 1,2-dithiolo-1,3-dithiole **169** and thietane-3-thione **170**. Both of these transformations are postulated to proceed *via* a spirothiirane intermediate **171**.

Ando and coworkers have extended their investigation of cumulene sulfurization to the cyclic biscumulene **172**, aiming to induce a transannular bridging reaction.¹⁸¹ Treating macrocycle **172** with an excess of sulfur in DMF at 120–130 °C for 10 hours yields cyclic polysulfides **173** (Scheme 37). In contrast, performing the same reaction in the presence of DBU affords cyclopentenethione **174** as the sole product. Single-crystal X-ray

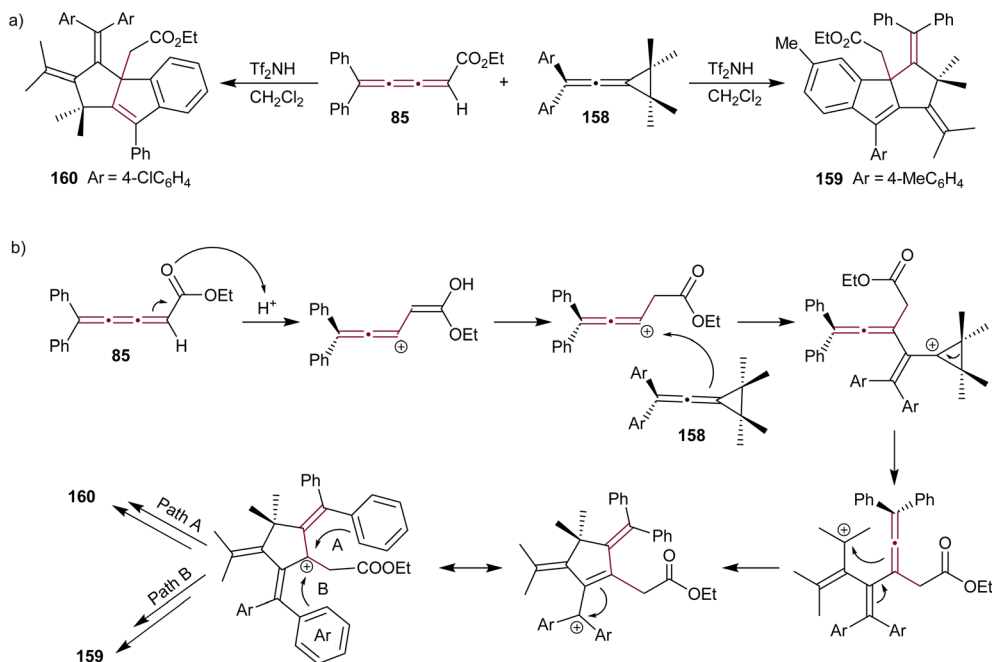




Scheme 33 (a) Synthesis of iodofulvenes **156** and **157** from [5]cumulene **155** and (b) the proposed mechanism.¹⁷⁶

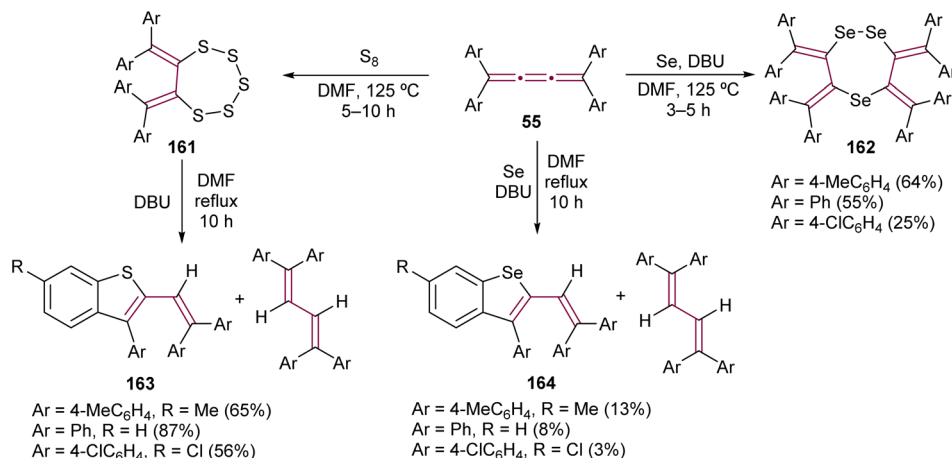
diffraction analysis unambiguously confirms the structures of both products. To identify whether **173** is the intermediate in the formation of **174**, the authors subjected polysulfide **173** to thermal decomposition under flow pyrolysis conditions, which affords the biphenylene derivative **175**. However, when

pentasulfide **173** is treated with DBU in DMF at room temperature, it produces the tetrasulfide **176**, whereas heating **173** to 160 °C generates the fused dithiolane **177** (Scheme 37). These observations indicate that the mechanism leading to the formation of **174** is fundamentally different from that for the



Scheme 34 (a) Reaction of [3]cumulene **85** with vinylidenecyclopropane **158** and (b) the proposed mechanism.¹⁷²



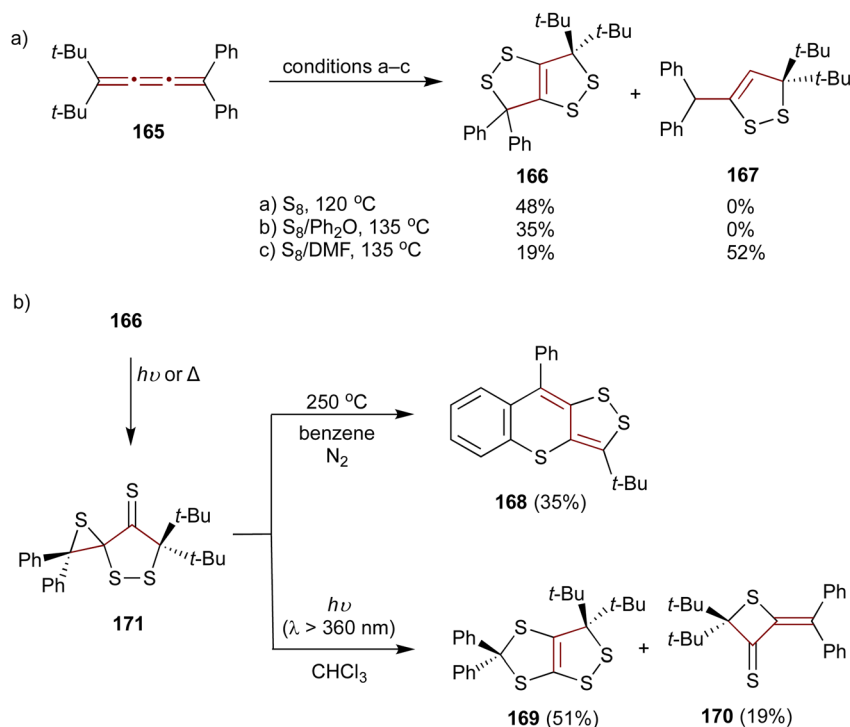
Scheme 35 Reactions of tetraaryl[3]cumulenes **55** with elemental sulfur and selenium.¹⁷⁸

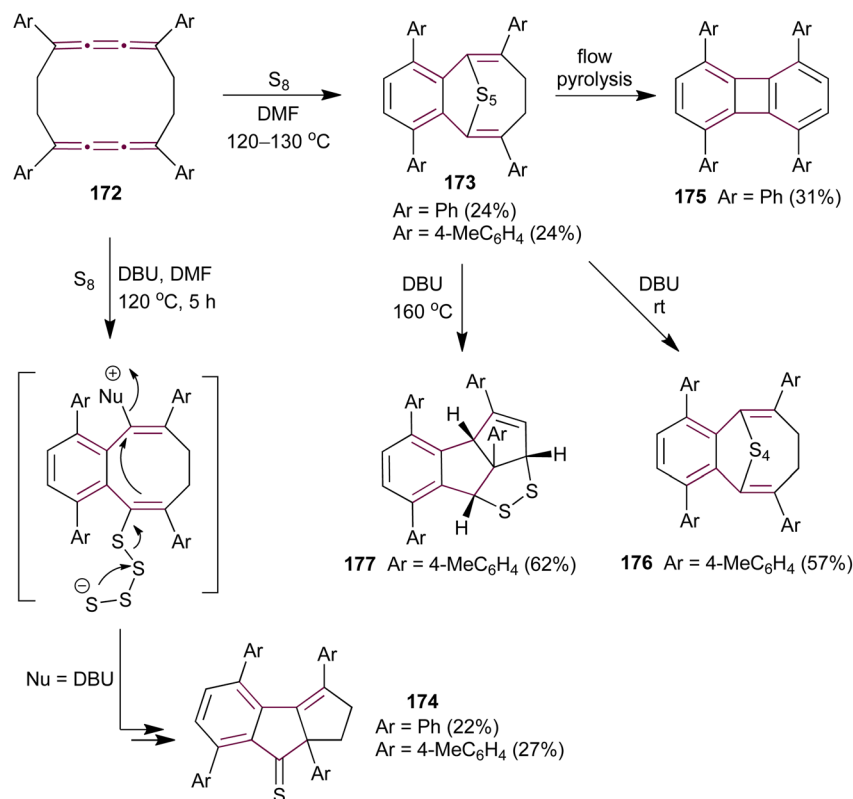
formation of **173** during the sulfurization of biscumulene **172**, presumably involving ionic sulfur species under basic conditions.

A study by Ando and coworkers in 1989 reports a convenient synthesis of episulfide [3]cumulenes **178a,b** and investigates their reactivity under thermal and photochemical conditions.¹⁸² Two transformations are predominant: conversion to thietane-thiones **179** and desulfurization to [4]cumulenes **180** (Scheme 38). Irradiation of the *tert*-butyl-substituted episulfide **178a** with UV light (300 nm) in chloroform at room temperature gives thietanethione **179a** and tetraene **180a** in 13% and 69% yields, respectively. When episulfide **178b** is heated in benzene under reflux, **179b** and **180b** are produced in 18% and 66% yields,

respectively. Mechanistic investigations suggest that the conversion of **178** to **179** proceeds *via* sulfur transfer involving diradical intermediates. Adding excess elemental sulfur during the thermolysis or photolysis of **178** fails to increase the yield of **179**. Moreover, the absence of dimerization products suggests that the bulky substituents of **178** create steric hindrance, favoring sulfur transfer as the principal reaction pathway.

The reaction of allenes with peracids yields products with high synthetic utility.¹⁸³ These reactions typically proceed through intermediate allene oxides (methyleneoxiranes), which then isomerize to form more stable cyclopropanones.¹⁸⁴ Extending this idea to the oxidation of cumulenes has also yielded some unique molecular architectures. For instance, the

Scheme 36 (a) Sulfurization of [3]cumulene **165** and (b) transformations of **166**.^{178–180}

Scheme 37 Sulfurization of biscumulene **172**.¹⁸¹

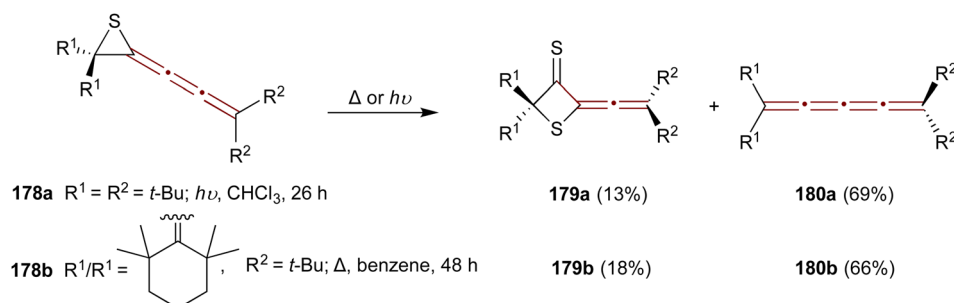
reaction of tetra(*tert*-butyl)[3]cumulene **181** with *m*-chloroperbenzoic acid (*m*-CPBA) leads to epoxidation of the terminal α -bond, presumably producing epoxide **182**, which readily isomerizes to cyclopropanone **183** (Scheme 39a).¹⁸⁵ Subsequent transformation of **183** with various reagents facilitates the synthesis of allenes, ketenes, and epoxyketones. In contrast, the oxidation tetraphenyl[3]cumulene **1** with *m*-CPBA yields tetralindione **184**, presumably *via* a double oxidation of the cumulene core followed by an intramolecular cyclization at the phenyl endgroup (Scheme 39b).¹⁸⁶

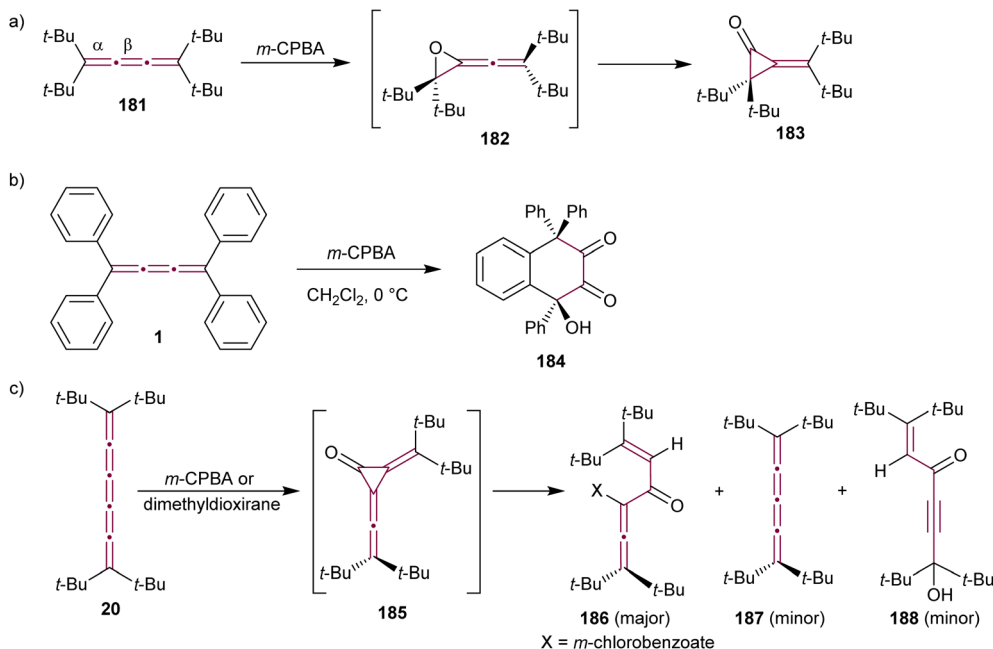
In 1992, Crandall and coworkers explored the oxidation of higher cumulenes (Scheme 39c).¹⁸⁷ Treatment of tetra(*tert*-butyl)[5]cumulene **20** with *m*-CPBA results in the transient formation of a cyclopropanone intermediate **185**, which subsequently transforms mainly to allenic ketone **186**. Minor quantities of

[4]cumulene **187** and ketone **188** are also detected in the reaction mixture.

The reaction of [3]cumulenes with vinylidene carbenes has been explored, similar to oxidation reactions, in relation to the reactivity of an allene.¹⁸⁸ The reaction of tetramethyl[3]cumulene **12** with the carbene generated from 1,1-dibromo-2-methylprop-1-ene mainly gives the expected adduct, [3]radialene **189**, together with a product of rearrangement, allene **190** (Scheme 40). Compounds **189** and **190** react further with the alkenyllithium carbene to produce the rather exotic radialene **191** and spirocyclopropane derivative **192**.

A study conducted by Skattebøl in 1965 examined the addition of dihalocarbenes to tetramethyl[3]cumulene **12** (Scheme 41a).¹⁸⁹ Addition of dichlorocarbene (generated from CHCl₃ and KO*t*Bu) results in a mixture of mono- and diadducts, **193** and **194** in 65% and 16% yields, respectively. In contrast, using

Scheme 38 Transformations of 1-episulfides **178** under thermal and photochemical conditions.¹⁸²



Scheme 39 Oxidation reactions of odd cumulenes: (a) **181**, (b) **1**, and (c) **20**.^{185–187}

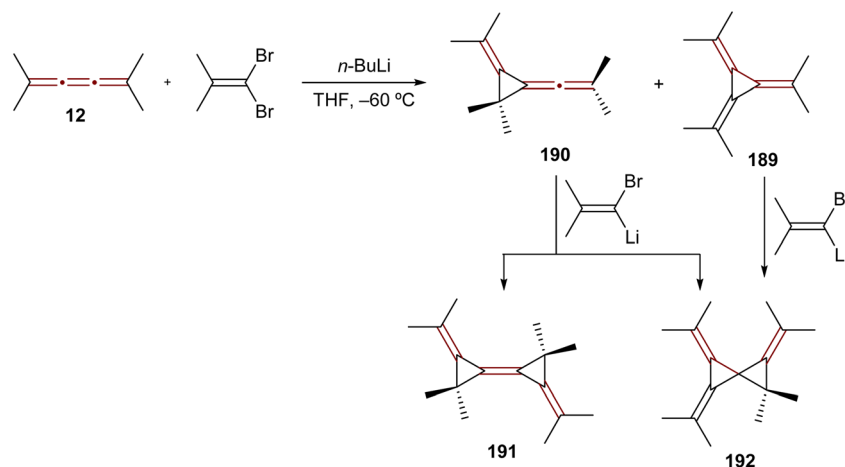
dibromocarbene (generated from CHBr_3 and $\text{KO}t\text{Bu}$) yields a complex mixture, from which only a small amount of an identifiable diadduct **195** is isolated. Treating **195** with methyl lithium at low temperatures gives tetramethyl[5]cumulene **22**, a species too unstable for detailed characterization and identified only by UV-vis spectroscopy, the structure of which has later been confirmed spectroscopically by Scott and DeCicco (Scheme 7b).⁵⁹

In 1974, Jochims and Karich described the first successful synthesis of [4]cumulene **196**, starting from the corresponding [3]cumulene **197**.¹⁹⁰ Their strategy involved dichlorocarbene addition to **197**, forming the intermediate **198**, which was subsequently converted into the [4]cumulene **196** (Scheme 41b). The structure of **196** was confirmed by elemental analysis and a range of spectroscopic methods. [4]Cumulene **196** is

surprisingly stable toward heat and oxygen, melting unchanged at 133 °C and showing no reactivity toward methyl lithium.

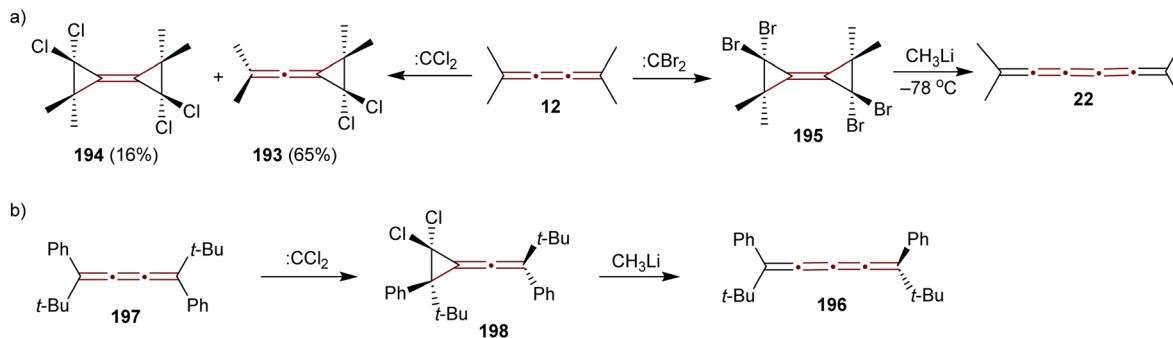
Reactive odd [*n*]cumulene intermediates

While the primary focus of this review is on the reactivity of stable, isolated cumulene species, it is important to acknowledge the broader reactivity landscape, which includes transient intermediates, strained cumulenes, and their organometallic variants. Although often unstable or highly specialized, these types of cumulenes play important roles in unique transformations and facilitate the construction of novel molecular architectures. Each represents a distinct and rich area of study, differing from classical cumulenes in bonding, stability, and reactivity. Although a detailed discussion of this area is beyond the scope of this review, a brief description is provided that



Scheme 40 Reaction of tetramethyl[3]cumulene **12** with a vinylidene carbene.¹⁸⁸





Scheme 41 Halocarbene addition to [3]cumulenes: (a) **194** and (b) **197**.^{189,190}

offers essential context for understanding the broader aspects of cumulene chemistry.

Strained cumulenes are known as reactive intermediates, with their intrinsic ring strain significantly enhancing their reactivity. These species engage in a spectrum of transformations, providing platforms for the construction of complex molecular architectures. Their unique reactivity is comprehensively reviewed by Johnson.¹⁹¹ Notable examples include medium-sized cyclic cumulenes such as 1,2,3-cyclo-decatriene, reported by Moore and Ozretich,¹⁹² as well as various other strained cyclic systems described by Johnson and co-workers.¹⁹³

Another notable example of a strained cumulene is the benzene isomer 1,2,3-cyclohexatriene,^{194–196} featuring cumulated double bonds but lacking aromatic stabilization. These structural features result in significantly higher free energy (+101 kcal mol⁻¹ relative to benzene) and enhanced reactivity.^{140,197} Recent studies by Garg and coworkers show that 1,2,3-cyclohexatriene, and its derivatives, participate in various strain-promoted reactions, including cycloadditions, nucleophilic additions, and σ -bond insertions (Scheme 42).^{197,198} For many organic chemists, 1,2,3-cyclohexatriene invites a comparison with *o*-benzyne, a strained aryne intermediate widely exploited for synthesis.^{199,200} Notably, complementary non-contact AFM and NMR studies have shown that benzyne might more accurately be described by a cumulenic bonding motif rather than the familiar ‘benzene-with-a-triple-bond’ depiction, a view that helps rationalize its characteristic reactivity.^{201,202}

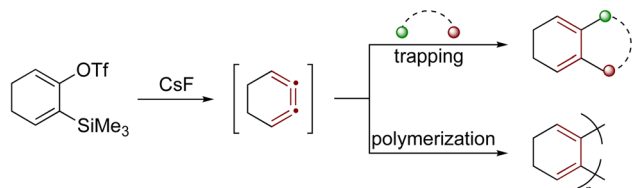
Reactive cumulene intermediates have also been exploited in polymer chemistry, where their cumulated π -systems confer high reactivity and enable selective chain-growth. Since 2020, Zhu and coworkers have been developing a family of copper-catalyzed polymerizations of propargylic substrates as

a promising strategy to access unique, all-carbon, alkyne-rich polymer backbones *via* reactive cumulene intermediates.²⁰³ Key examples are summarized in Scheme 43 and illustrate how transient cumulene intermediates enable the creation of diverse polymer architectures.^{204–208} Representative examples include the dehydrative polymerization of propargylic alcohols,²⁰⁵ copolymerization of propargyl carbonates with aryldiazomethanes,²⁰⁸ stepwise-chain growth condensation routes to conjugated [5]cumulene polymers,²⁰⁴ polymerization *via* catalytic element-cupration,²⁰⁶ and ring-opening transformations of cyclic carbonates.²⁰⁷ By tuning propargylic substrates such as carbonates, alcohols, or acetates along with reaction conditions and external nucleophiles, a wide variety of polymeric alkynes and cumulene-containing frameworks have been realized.

Organometallic chemistry of cumulenes encompasses both short-lived cumulene-containing intermediates and isolable complexes (including metallacyclocumulenes), forming a substantial field in its own right. Coordination with transition metals can stabilize and polarize cumulenic frameworks, unlocking distinctive reactivity relevant to synthesis and catalysis, including selective C–C activation, migratory insertion/fragmentation sequences, and cross-coupling pathways. As this area has been comprehensively surveyed, we refer readers to reviews for a detailed coverage of bonding models, preparative strategies, and representative applications.^{209,210}

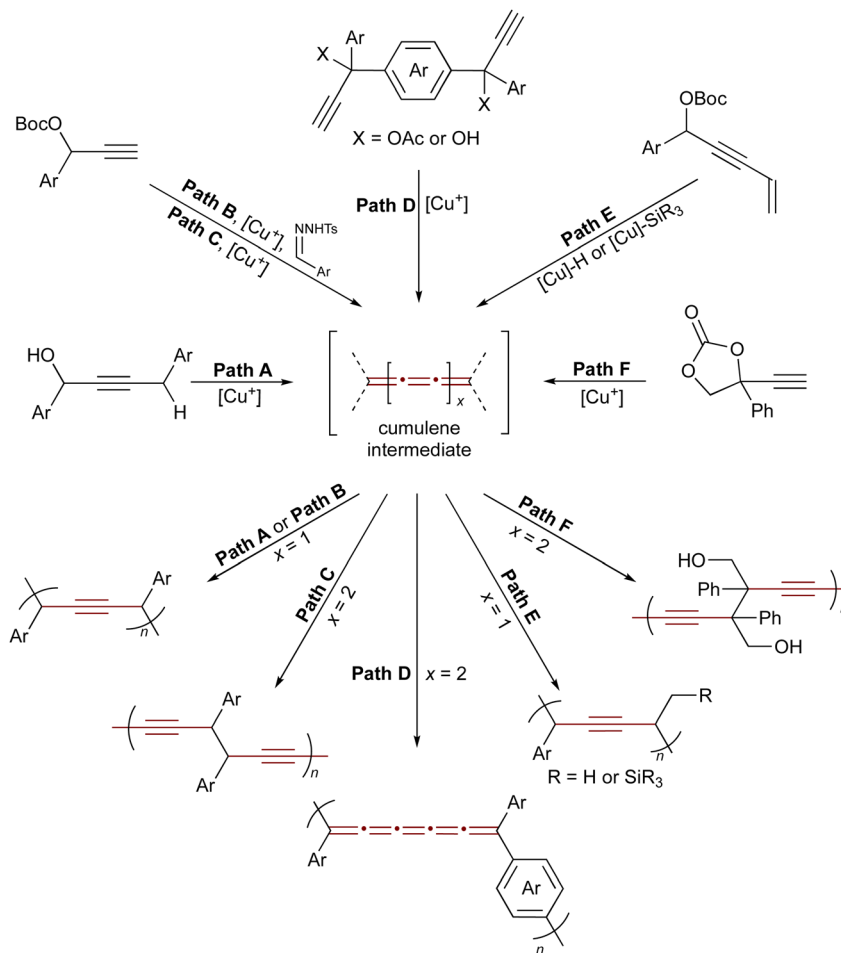
Even cumulenes

Even [*n*]cumulenes (*n* = 4, 6, 8...) feature a profoundly different π -electron system compared to odd [*n*]cumulenes (*n* = 3, 5, 7...) due to the presence of two degenerate, orthogonal π -electron systems (see Fig. 1).²¹¹ A close structural relative to even [*n*]cumulenes, allenes, are valuable precursors in modern organic synthesis due to their π -electron-rich systems and intrinsic axial chirality (depending on the substitution pattern of endgroups). The reactivity and chemistry of allenes have been discussed in numerous reviews^{212–215} and will not be revisited here. Extending the chemistry of allenes to longer even [*n*]cumulenes has been sporadic, likely due to the inherent instability of these unsaturated substrates. To date, [4]cumulenes represent the longest even [*n*]cumulenes successfully isolated; a single example of *in situ* formation of a [6]cumulene has also been reported (*vide infra*).^{216,217} With the development of endgroups designed to improve stability,



Scheme 42 Reactivity of the strained cumulene intermediate: 1,2,3-cyclohexatriene.¹⁴





Scheme 43 Copper-catalyzed polymerization of propargylic compounds. Path A: dehydrative polymerization of propargylic alcohols.²⁰⁵ Path B: copolymerization of propargyl carbonates and aryldiazomethanes.²⁰⁸ Paths C and D: synthesis of polydiynes *via* a dimerization and polymerization sequence of propargylic electrophiles.²⁰⁴ Path E: condensation polymerization enabled by the 1,2-regioselective hydro- and silylcupration of enyne-type propargylic electrophiles.²⁰⁶ Path F: ring-opening polymerization of cyclic carbonates.²⁰⁷

the breadth of isolable even $[n]$ cumulenes has been slowly expanded, allowing for the study of their reactivity. Key transformations of even $[n]$ cumulenes are similar to those of odd $[n]$ cumulenes and include cyclooligomerization, cycloaddition, reactions with metals, hydrogenation, halogenation, oxidation, as well as other miscellaneous reactions encountered during their synthesis.

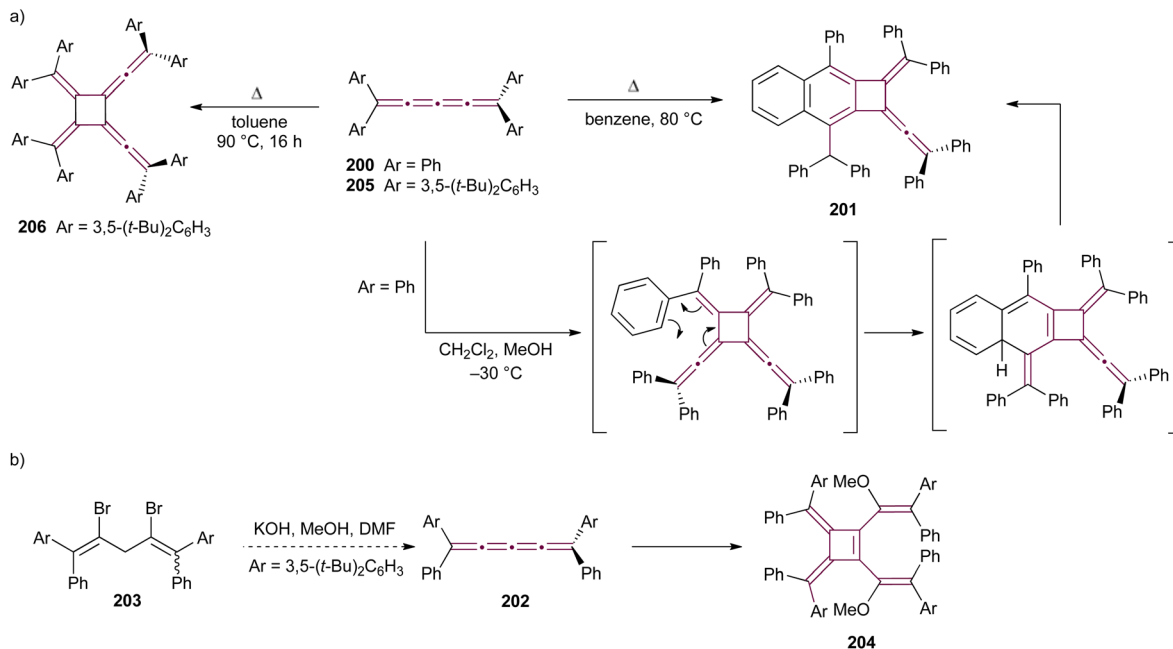
Cyclodimerization

Shortly after Shechter and Tiers confirmed the structure of the dimerization product from tetraphenyl[3]cumulene **1**, originally reported by Brandt in 1921,^{8,43} the striking photostability of tetraphenyl[4]cumulene **200** was highlighted by Fischer and Fischer.²¹⁸ They attributed this behavior to two independent molecular π -orbitals in the [4]cumulene framework and the correspondingly larger HOMO–LUMO gap characteristic of even $[n]$ cumulenes. The synthesis of **200** was independently achieved by Kuhn,²¹⁹ Fischer,²¹⁸ and then Ratts,²²⁰ although their observations differed. While Kuhn did not report dimerization, both Fischer and Ratts suggested that **200** undergoes thermal dimerization. Unlike **1**, which converts to radialene **6**, Fisher

noted that the product from dimerization of **200** shows one non-aromatic proton signal in its ¹H NMR spectra. This analysis suggests that dimerization of **200** to form a radialene without incorporation of the aryl endgroups is unlikely. Although the exact structure of the product remained unclear, the formation of a naphthalene-containing moiety was proposed by Fischer (*vide infra*). In 2009, the structure of the product was proposed by Iyoda as the dimer 1,2-dihydrocyclobuta[*b*]naphthalene **201**, formed *via* thermal dimerization of **200** in benzene under reflux to give an intermediate, which then underwent a 6 π -electrocyclization followed by a proposed H-shift (Scheme 44a).²²¹ In a subsequent synthetic report by Tykwinski and coworkers in 2023, the [4]cumulene **200** was isolated in *ca.* 42% yield and 95% purity. Attempts to further purify **200** through recrystallization led to the formation of **201** (Scheme 44a), which was confirmed through spectroscopic methods as well as X-ray crystallography.²²¹ Interestingly, the [2 + 2] dimerization of **200** occurred in a head-to-head manner, similar to the syn-orientation determined by Stang and coworkers for odd cumulenes.²²²

Increasing the steric bulk of terminal endgroups has been investigated as a strategy to suppress the intermolecular

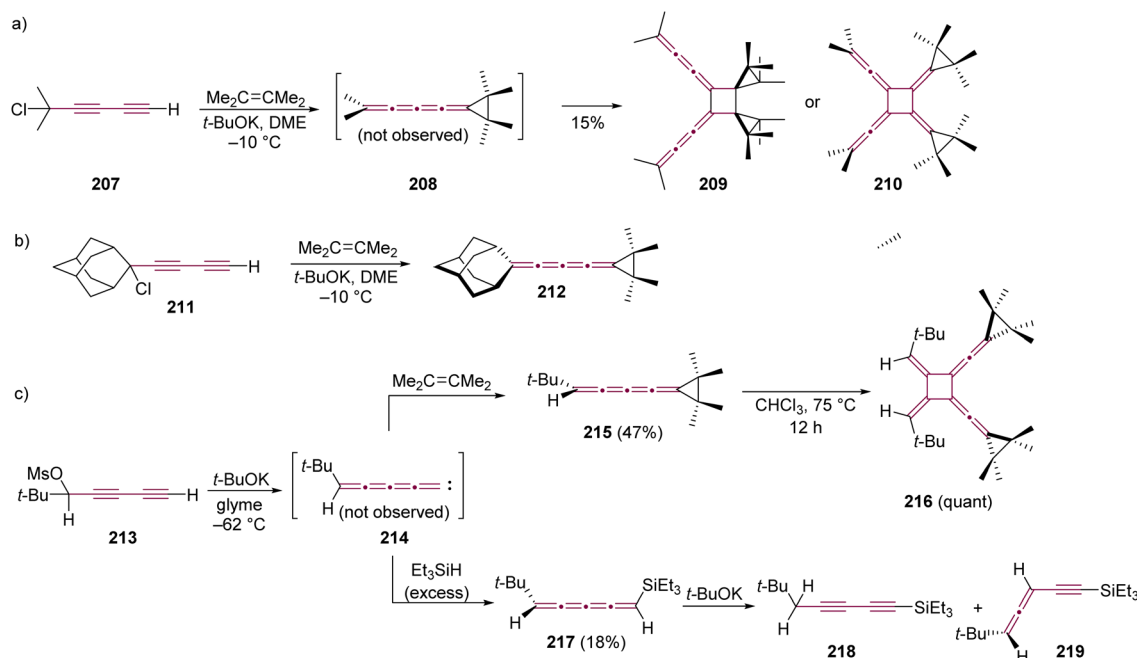




Scheme 44 (a) Intermolecular dimerization of tetraphenyl[4]cumulene **200** and tetra(3,5-di-*tert*-butyl)phenyl[4]cumulene **205**. (b) Formation of dimer **204** during the attempted synthesis of [4]cumulene **202**.^{211,220,221}

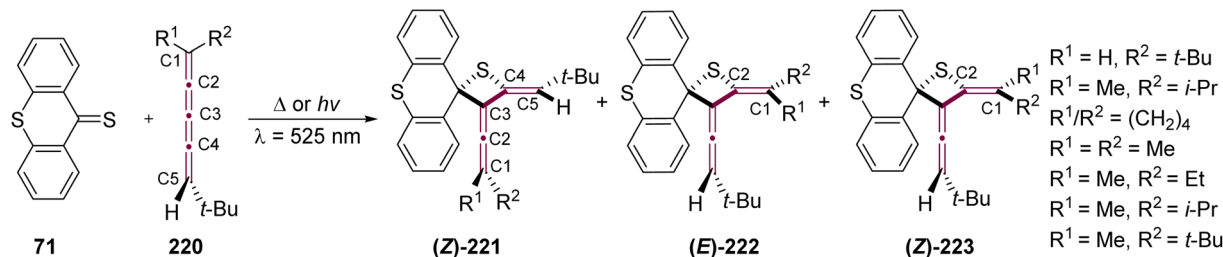
dimerization pathway, utilizing the sterically bulky (3,5-di-*tert*-butyl)phenyl group. The synthesis of [4]cumulene **202**, bearing a combination of phenyl and (3,5-di-*tert*-butyl)phenyl endgroups, has been attempted through the exhaustive elimination of HBr from the 1,4-diene **203** (Scheme 44b). However, this approach yielded a complex mixture of products, and the formation of **202** was not observed. From this mixture, the dimeric product **204** was isolated and identified by X-ray

crystallography. It remains unclear whether **204** formed through trapping of an intermediate during the elimination of **203** or during purification of the crude reaction mixture. The instability of **202**, compared to the isolable **200**, is attributed to the increased electron density introduced by the (3,5-di-*tert*-butyl)phenyl endgroups, which offsets the potential stabilization provided by steric shielding.



Scheme 45 Synthesis and subsequent reactions of [4]cumulenes based on trapping of vinylidene carbenes bearing (a) methyl **207**, (b) adamantyl **211**, and (c) *t*-Bu **213** endgroups.^{224–226}





Scheme 46 Thermal and photochemical [2 + 2] cycloaddition of thioxanthenethione **71** with [4]cumulenes **220**.²²⁸

Since two (3,5-di-*tert*-butyl)phenyl endgroups are insufficient for stabilization, [4]cumulene **205**, bearing four bulky substituents, has been targeted and synthesized. Cumulene **205** dimerizes when heating a sample to 90 °C in dry toluene, as revealed by ¹H NMR spectroscopy. Similar to the thermal dimerization of **200**, the dimerization of **205** result in the formation of a head-to-head dimer **206**, as confirmed by X-ray crystallographic analysis (Scheme 44a), albeit in only trace amounts.

In the 1960s, extended unsaturated carbenes, such as alkylidenes and vinylidenes, were extensively studied through the trapping of reaction intermediates.²²³ For example, Stang and coworkers have investigated the reactivity of vinylidene carbenes through the addition reactions to olefins, yielding cumulenic products. From both mechanistic and synthetic perspectives, carbene trapping has emerged as a valuable approach for advancing the understanding of cumulene chemistry.^{66,222}

Le Noble and coworkers have explored a carbene trapping method to form an even [4]cumulene, through base-induced elimination of the diyne precursor **207** and tetramethylethene as the trapping agent (Scheme 45a).^{224,225} The [4]cumulene **208** could not be isolated, and a dimer was obtained as the only product. The head-to-head dimerization was proposed based on earlier findings by Stang and coworkers,²²² and among the two possible isomers (**209** and **210**), the former was deemed more consistent with the spectroscopic data. Interestingly, replacing the methyl endgroups with a bulkier adamantylidene group in a similar carbene trapping experiment with **211** yielded the stable [4]cumulene **212**, which did not undergo cyclo-oligomerization (Scheme 45b).²²⁵

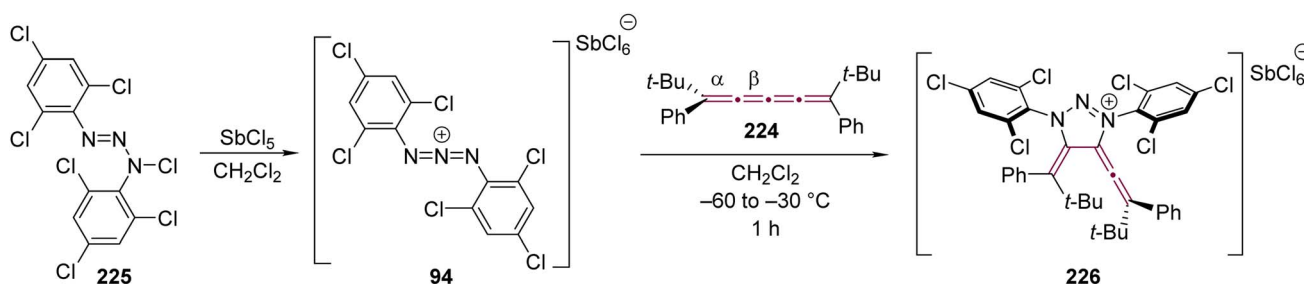
Stang and Learned, in 1988, reported on the reactivity of a vinylidene carbene system terminated at one end with a single *tert*-butyl group (Scheme 45c).^{226,227} Starting from diyne **213**,

elimination at a low temperature (−62 °C) produces carbene **214**, which is trapped with tetramethylethene to afford the [4]cumulene **215** in 47% yield. Cumulene **215** is relatively unstable and undergoes cyclodimerization quantitatively to form **216** after several days at room temperature or a few hours upon heating in chloroform.²²⁷ Alternatively, trapping **214** with Et₃SiH yields the unstable [4]cumulene **217** in 18% yield. In the presence of *t*-BuOK, **217** undergoes a prototropic rearrangement to give a mixture of diyne **218** and enyne **219** (Scheme 45c).

Cycloaddition reactions

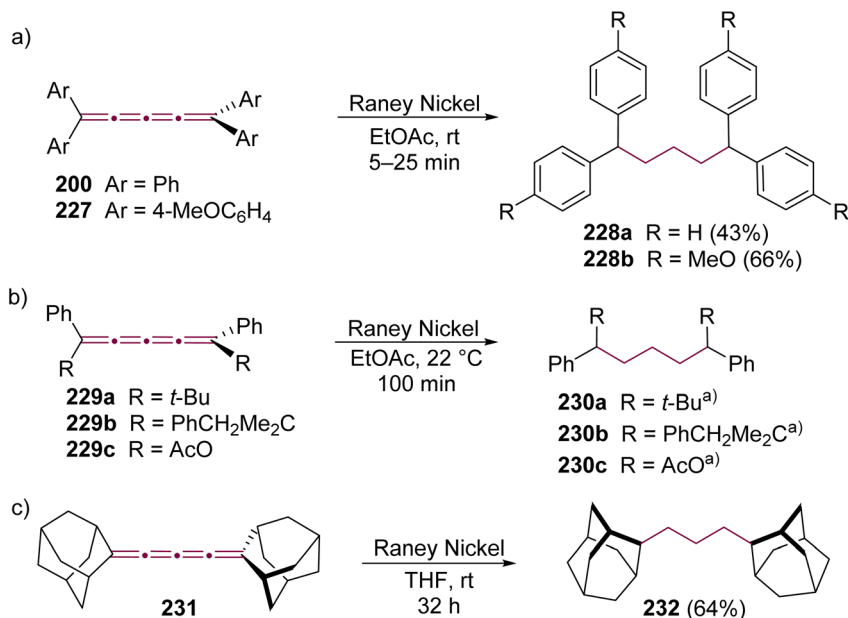
Visser and coworkers have studied a series of photochemical and thermal [2 + 2] cycloaddition reactions between thioxanthenethione **71** and [4]cumulenes **220** (Scheme 46).²²⁸ The thermal reactions were conducted using a solution of **71** (0.02 M) and **220** at room temperature and were typically completed within 15 minutes. To prevent competition from the rapid thermal reaction, the photochemical reactions were performed in dilute solutions of **71** (0.005 M) at −70 °C. [4]cumulenes **220** bearing a range of endgroups and substitution patterns were investigated, and the reactions consistently produced three isomeric thietanes: (*Z*)-**221**, (*E*)-**222**, and (*Z*)-**223**. Notably, in contrast to earlier photochemical studies on odd [*n*]cumulenes that show reactivity at C1 and C2, the thio-carbonyl functional group in its triplet state preferentially attacks position C2 or C4 of the [4]cumulenes. Interestingly, the size of the substituent does not significantly affect the product ratios under photochemical conditions. Under thermal conditions, however, steric hindrance from the substituents plays a dominant role in determining the product distribution.²²⁸

In 1997, a [3 + 2] cycloaddition reaction was reported by Jochims and coworkers using [4]cumulene **224** and the 1,3-diaza-2-azoniaallene salt **94**, formed *in situ* by reacting



Scheme 47 Reaction of [4]cumulene **224** with 1,3-diaza-2-azoniaallene salt **94**.¹¹⁸





Scheme 48 Hydrogenation reaction with RANEY® on [4]cumulenes: (a) **200** and **227**, (b) **229a–c**, and (c) **231**; ^{a)} – yield not given.^{219,229,230}

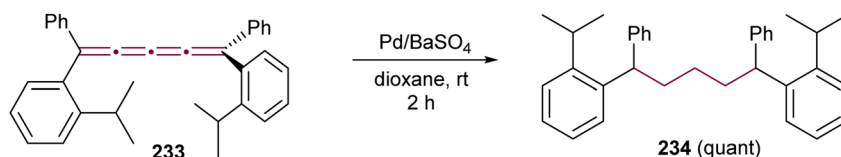
antimony pentachloride with *N*-chlorotriazine **225** (Scheme 47).¹¹⁸ The salt **94** was added to a solution of [4]cumulene **224** in CH₂Cl₂, producing a red suspension *via* the [3 + 2] cycloaddition reaction and, upon purification, the triazolium salt **226** was isolated. X-ray crystallographic analysis revealed that the cycloaddition proceeded regioselectively at one of the β-bonds of the [4]cumulene moiety. The formation of triazolium salt **226** was calculated to be exothermic by 40 kcal mol⁻¹, which is markedly higher than 18 kcal mol⁻¹ for the corresponding cycloaddition reaction of **94** with analogous [3]cumulene **95**. This finding suggests a more thermodynamically favorable product for the reaction with [4]cumulenes, although the overall reactivity would also depend on the activation barrier, which is not explicitly addressed in this study. A comparison of the reactivity of salt **94** with both [4]cumulene **224** and [3]cumulene **95** (Scheme 20) raises intriguing questions regarding the steric and electronic factors that govern cycloaddition regioselectivity in cumulenes. Further investigations into alternative reaction conditions and substituent effects may provide a more comprehensive understanding of the reactivity of cumulenes.

Organometallic reactions

Before the advent of advanced spectroscopic techniques, hydrogenation was widely employed to identify the presence and number of non-aromatic double bonds in cumulenes.²¹⁹ [4]cumulenes, which contain four contiguous double bonds,

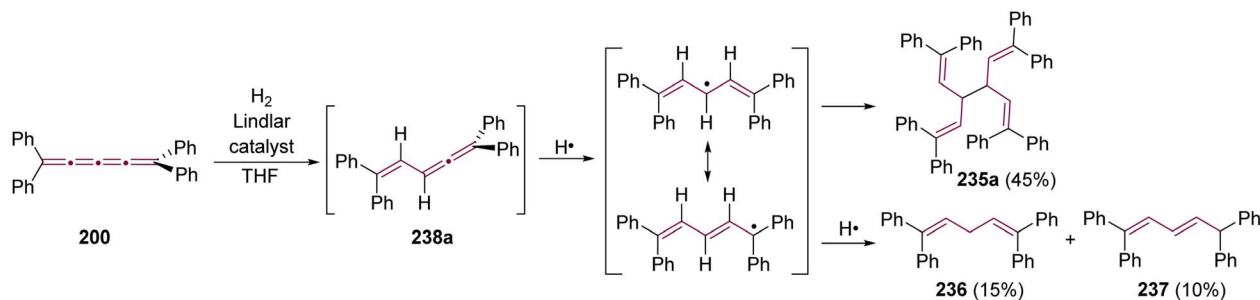
can undergo complete hydrogenation to carbon–carbon single bonds with consumption of four molar equivalents of hydrogen.^{219,229} In most of this section, we discuss the hydrogenation of [4]cumulenes using RANEY® and palladium catalysts to confirm their constitutional molecular formulas. We also highlight the correlation between the reaction time and the impact of various endgroups on the hydrogenation process. Finally, we examine the partial hydrogenation of [4]cumulenes with Lindlar's catalyst and aluminum amalgam, illustrating the formation of intriguing side products.

To the best of our knowledge, the first hydrogenation of even [*n*]cumulenes featuring aryl endgroups (**200** and **227**) was reported in 1964 by Kuhn and Fischer.²¹⁹ When these [4]cumulenes are treated with RANEY® in ethyl acetate, approximately four equivalents of hydrogen are consumed in each case, yielding tetraarylpentanes **228a** and **228b** in 43 and 66% yields, respectively, within 5–25 minutes (Scheme 48a). Subsequently, in 1977, Karich and Jochims investigated the reduction of [4]cumulenes bearing aryl and bulky alkyl endgroups (**229a,b**, Scheme 48b),²²⁹ while in 1986 Nader and Brecht examined a [4]cumulene bearing both aryl and carboxylate endgroups (**229c**, Scheme 48b).²³⁰ In all these cases, the results mirror those of Kuhn and Fischer, with [4]cumulenes **229a**,¹⁹⁰ **229b**,²²⁹ and **229c**²³⁰ each consuming approximately four equivalents of H₂ to produce the corresponding 1,1,5,5-tetrasubstituted pentanes **230a–c**, albeit the time necessary for full hydrogenation is



Scheme 49 Hydrogenation of [4]cumulene **233** using Pd/BaSO₄.²³¹



Scheme 50 Hydrogenation of **200** with Lindlar's catalyst.²¹⁸

extended to 30–100 min (Scheme 48b). A comparison of these studies suggests that bulky alkyl endgroups likely shield the cumulene core, thereby slowing the hydrogenation reactions.

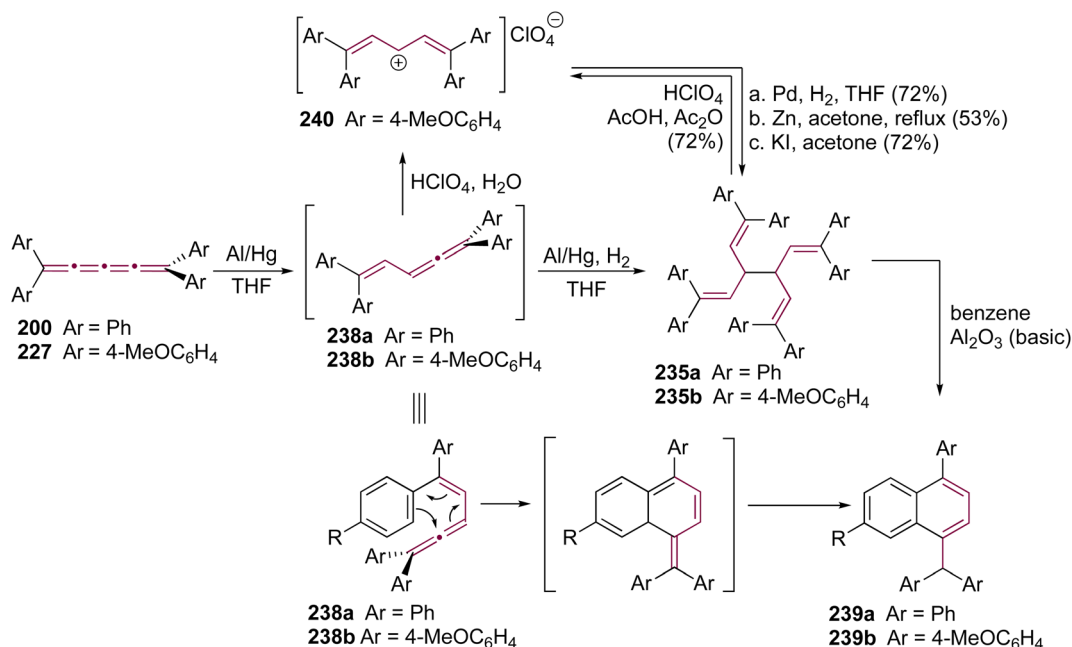
Karich and Jochims described the hydrogenation of adamantylidene-encapped [4]cumulene **231**, demonstrating the effect of fully replacing aryl endgroups with an alkyl substituent (Scheme 48c).²²⁹ They observed that the reactivity of **231** decreases significantly, with the reaction requiring 32 hours to yield the fully hydrogenated product **232**. The authors attributed the reduced reactivity, in part, to the low solubility of **231** in THF, which may have hindered the reaction speed. If one considers the progression from **200/227** (5–25 min) to **229** (100 min) to **231** (32 h), however, it appears equally plausible that the alkyl substitution itself contributes to the slower reaction rate.

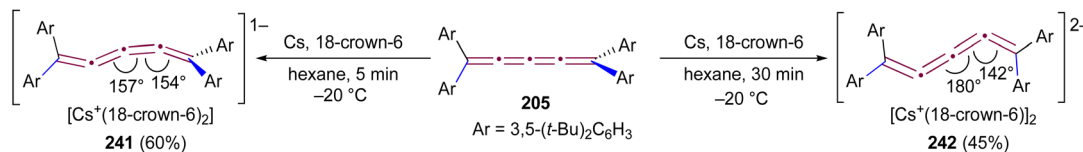
In 1977, von Herrath and Rewicki demonstrated the hydrogenation of cumulene **233** to 1,1,5,5-tetrasubstituted pentane **234** using Pd/BaSO₄ (Scheme 49).²³¹ Compared to tetraaryl[4]cumulenes **200** and **227**, which are fully hydrogenated within 25 min, the reaction with Pd/BaSO₄

requires two hours to fully convert the product, also with a quantitative yield.

In 1964, Fischer and Fischer reported a partial hydrogenation of **200** using Lindlar's catalyst.²¹⁸ The reaction consumes approximately 1.9 equivalents of H₂, rather than the four equivalents required for full hydrogenation using Pd/BaSO₄ or RANEY®. Under these conditions, the major product arises from reductive dimerization to yield **235a** (45% yield), accompanied by the minor diene products **236** and **237** in 15% and 10% yields, respectively (Scheme 50). The authors proposed a mechanism involving the initial formation of intermediate **238a**, which then undergoes addition of an H[•] radical to produce an intermediate radical species. This radical preferentially dimerizes to form **235a**, while a small proportion undergoes H[•] abstraction, leading to the formation of dienes **236** and **237** (Scheme 50).

Fischer has explored the species involved in the reduction of **200** and **227** using Al/Hg amalgam (Scheme 51).²¹⁸ Attempts to isolate the intermediates **238a** and **238b** from Al/Hg reduction *via* column chromatography (Al₂O₃) instead yield the colorless

Scheme 51 Hydrogenation reaction of [4]cumulenes **200** and **227** with Al/Hg amalgam.²¹⁸ Oxidation of **238b** with HClO₄ and subsequent dimerization to **235b**.²³² Proposed mechanism for the formation of naphthalenes **239a,b**.²¹⁸



Scheme 52 Chemical reduction of [4]cumulene **205** with cesium metal to form mono- and dianions (bonds for calculating the torsion angle Φ are shown in blue).²³⁴

naphthalene derivatives **239a** and **239b** (Scheme 51). These compounds are postulated to arise from triene intermediates **238a** and **238b**, respectively, which undergo a 6π -electrocyclization followed by a proposed [1,7]-hydride shift. This transformation closely resembles the mechanism previously proposed for the formation of **201** from [4]cumulene **200** (Scheme 44).

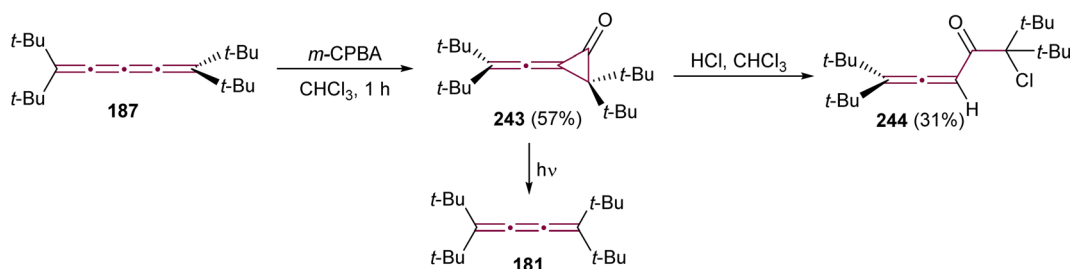
On the other hand, Wizinger and coworkers reported that treating **238b** with perchloric acid produces the oxidized species **240** as a blue solid.^{232,233} Subsequent reduction of **240** yields the dimer **235b** as the predominant product regardless of the conditions employed. This transformation resembles the formation of **235a** from **238a** previously reported by Fischer using Lindlar's catalyst (Scheme 50).²¹⁸ Notably, the formation of the dimer **235b** is reversible, *i.e.*, treating **235b** with perchloric acid in acetic acid and acetic anhydride regenerates the oxidized salt **240** in 72% yield.

The first successful isolation of the reduced species formed from a [4]cumulene was reported by Petrukhina, Tykwinski, and coworkers in 2024. The authors explored the redox characteristics of **205** using cesium metal as a reducing agent and observed the formation of both a monoanionic radical and a dianionic species.²³⁴ The authors noted that the reduced species obtained from the reaction of **205** with either sodium or potassium metal is too reactive for isolation. Crystallographic data reveal that, upon stepwise reduction, the [4]cumulene undergoes significant geometric changes that include twisting of the endgroups from an orthogonal geometry in **205** toward planarization in mono- and dianions **241** and **242**, respectively (Scheme 52). Using X-ray structural analysis, the torsion angle between opposing endgroups (blue bonds in Scheme 52) is reduced from $\Phi = 72^\circ$ in neutral **205** to $\Phi = 42^\circ$ in the monoanion **241** to $\Phi = 0^\circ$ (*i.e.*, coplanar) in the dianion **242**. These findings suggest important implications for the electron-handling capabilities of cumulenes, particularly in the context of redox-active materials.

Miscellaneous reactions

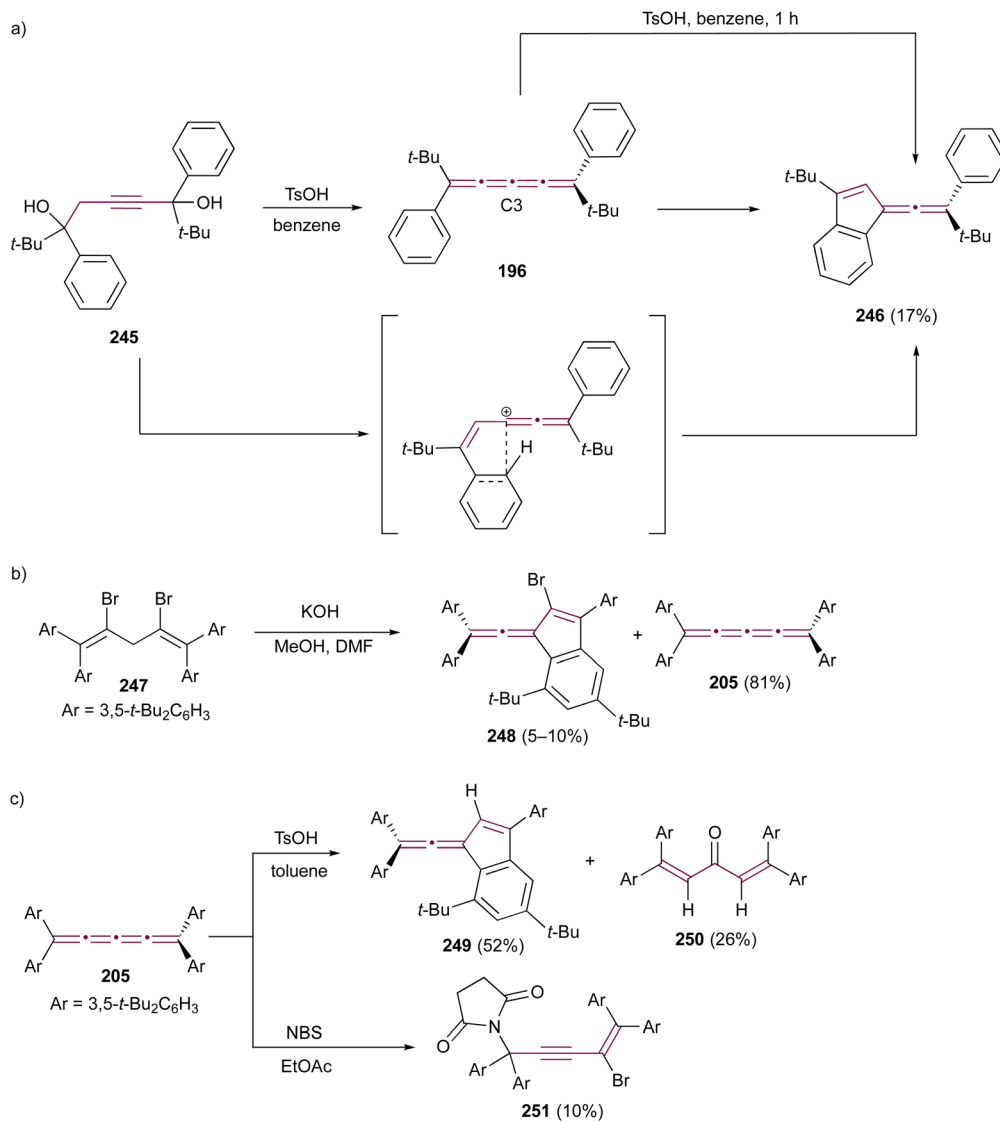
Crandall and coworkers have examined the reactivity of [n]cumulenes with *m*-CPBA and dimethyldioxirane (DMDO), proposing that the initially formed (not isolated) epoxide rapidly isomerizes to a relatively stable cyclopropenone intermediate. In one example, the reaction of tetra(*tert*-butyl)[4]cumulene **187** with one equivalent of *m*-CPBA yields the cyclopropenone intermediate **243** (Scheme 53).¹⁸⁷ Over time, exposure of **243** to sunlight induces slow decarbonylation, affording the [3]cumulene **181**. Alternatively, treatment of **243** with hydrochloric acid furnishes the allenyl ketone **244**.

Nakagawa and coworkers reported the synthesis of [4]cumulene **196** through dehydration of diol **245** using *p*-toluenesulfonic acid in benzene under reflux (Scheme 54a). The resulting yellow crystalline solid was initially proposed to be **196**. However, subsequent hydrogenation experiments revealed the reduction of only three double bonds, inconsistent with the proposed structure of **196**.²³⁵ In 1962, Kuhn and Schulz independently reported the synthesis of indene **246**,²³⁶ which showed hydrogen consumption and spectral data consistent with the molecule described by Nakagawa. This finding strongly suggests that a structural revision is needed for Nakagawa's proposed product. While the hydrogen consumption and spectra of **246** aligned with Nakagawa's final product, the revision depended primarily on indirect data, given that no [4]cumulenes were yet known at the time. Attempts by Kuhn and Schulz to prepare [4]cumulene **196** in 1963 were unsuccessful.^{237,238} A breakthrough was noted in 1974, when Jochims and Karich successfully synthesized [4]cumulene **196** *via* carbene insertion followed by dehalogenation, marking the first successful isolation of this species.^{190,229} In the same study, treating **196** with TsOH triggered intramolecular cyclization at the *ortho*-position of the phenyl endgroup and C3 of the cumulene chain (Scheme 54a).¹⁹⁰ The resulting yellow crystalline solid was characterized



Scheme 53 Oxidation of [4]cumulene **187** with *m*-CPBA.¹⁸⁷





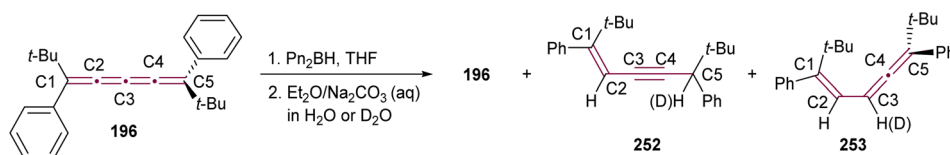
Scheme 54 Reactions of (a) [4]cumulene **196** with TsOH, (b) formation of **248**, and (c) reactions of [4]cumulene **205** with TsOH or NBS.^{190,211,235}

as the indene structure **246**, consistent with the findings of Kuhn and Schulz.

Indene formation was documented in 2023 for the sterically encumbered [4]cumulene **205** when treated with TsOH in toluene.²¹¹ During the synthesis of **205**, the elimination reaction of divinylbromide **247** generated both the desired [4]cumulene **205** and the brominated indenylidene **248** (Scheme 54b). Because it is difficult to envision a plausible mechanism for formation of **248** under the strongly basic conditions used to eliminate HBr from **247**, the authors proposed that **248** is most

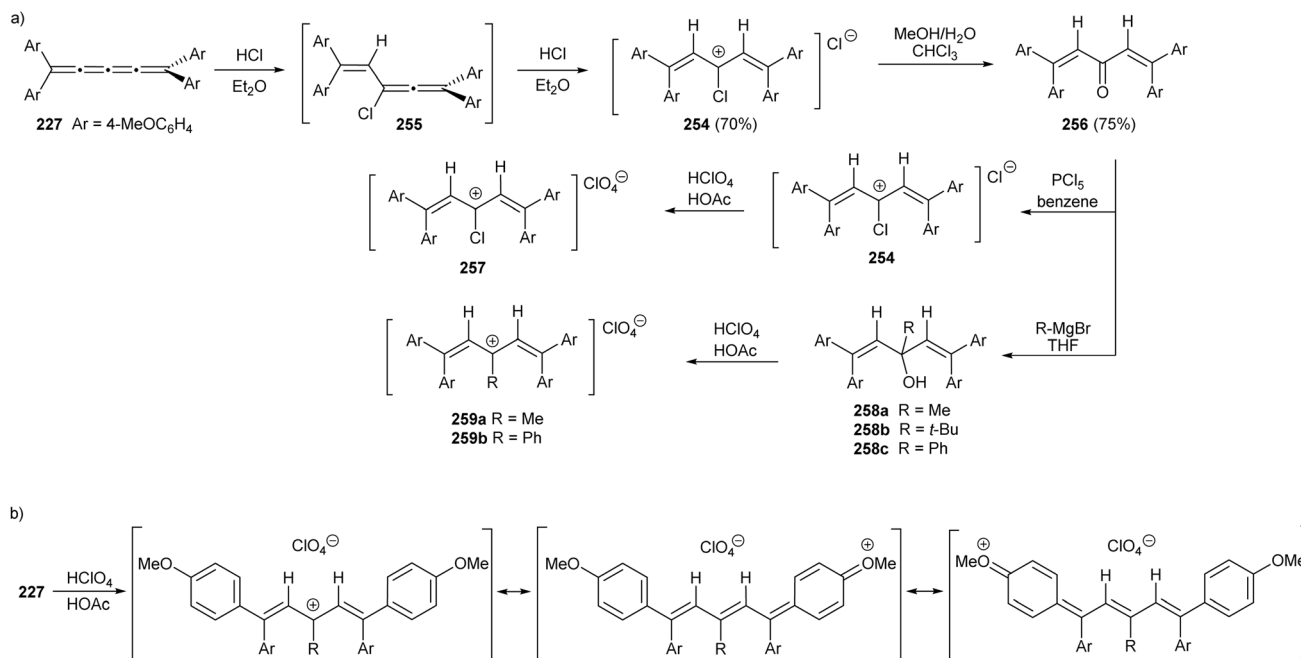
likely produced by bromination of **205** during the workup process.

In analogy to the formation of **246** from **196**, subjecting [4]cumulene **205** to TsOH induced intramolecular cyclization, giving **249**, along with **250**, which arose from the acid-catalyzed hydration of **205** (Scheme 54c). Notably, the presence of acid is essential, as stirring **205** in a mixture of acetone and water (1 : 1) at room temperature did not yield either **249** or **250**. Although not explored in detail, an attempt to brominate **205** with the mild brominating agent NBS did not reportedly provide **248**;



Scheme 55 Hydroboration of [4]cumulene **196** with Pn_2BH ($\text{Pn} = (+)\text{-}\alpha\text{-pinene}$).²³⁹





Scheme 56 (a) Reaction of [4]cumulene **227** with HCl and (b) resonance stabilization of the salt **259** proposed by Fischer.²¹⁸

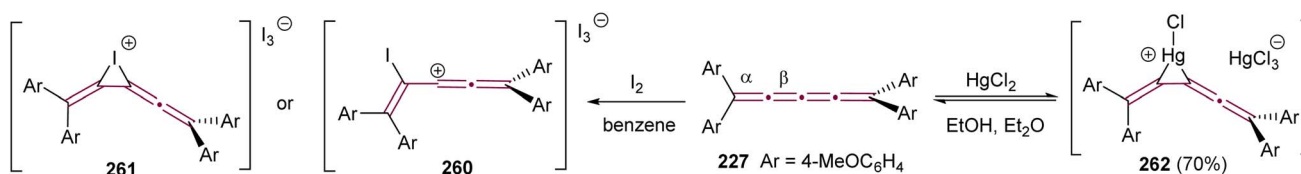
instead, **251** was obtained and characterized by X-ray crystallography.

Jochims and coworkers have reported the partial hydroboration of [4]cumulene **196** with di(3-pinanyl)borane (Pn₂BH), toward enantioselective enrichment *via* kinetic resolution. Treating **196** with a substoichiometric amount of Pn₂BH (0.67 equivalents) returns 33% of unreacted **196** with approximately 5% enantioenrichment, presenting a potential approach to enantiomeric resolution of the [4]cumulene scaffold. Enyne **252** is isolated as the major product and triene **253** as the minor product (Scheme 55).²³⁹ Deuteration experiments reveal that boronation occurs at C3 of the cumulene core, while hydride is delivered to C2. Subsequent quenching with H₂O or D₂O introduces a proton or deuterium at C5 (yielding **252**) or C3 (yielding **253**).

Tetraphenyl[4]cumulene **200** reacts in the presence of either hydrochloric or formic acid, but the hydrolyzed product is not sufficiently stable for characterization (not shown). Fischer and Fischer report that treating [4]cumulene **227** with two equivalents of HCl yields black-violet crystals identified as the pentadienyl salt **254** (Scheme 56a).²¹⁸ They hypothesized that the reaction proceeds *via* a pentatriene intermediate **255**, formed by the addition of one equivalent of HCl to [4]cumulene **227**,

although **255** is not isolated. Subsequent treatment of pentadienyl salt **254** in chloroform with a mixture of methanol and water (9 : 1) gives colorless crystals of pentadieneone **256** in 75% yield. Treating **256** with PCl₅ in benzene gives a deep violet solution whose UV-vis spectrum matched that of **254**, and its further reaction with perchloric acid produces the brown-violet perchlorate salt **257**. Alternatively, the reaction of pentadieneone **256** with various Grignard reagents yields alcohols **258a-c**, which subsequently form the corresponding perchlorate salts **259a,b** upon reaction with perchloric acid. The authors suggested that the phenyl endcapped derivatives lack the resonance stabilization provided by anisole endgroups, thus accounting for the reduced stability (Scheme 56b).

Treating tetraanisyl[4]cumulene **227** with iodine yields a gray-black crystalline precipitate proposed to be the allenic cation **260** (Scheme 57). Iodination at the β-bond is suggested on the basis of IR and UV-vis analyses, as well as comparisons with results from the protonation reactions (*vide supra*). Although not explicitly mentioned in the study, it is worth considering that the structure takes the form of an iodonium intermediate **261**.²⁴⁰ Intermediates or products from the analogous halogenation reactions of [4]cumulene **227** using Br₂

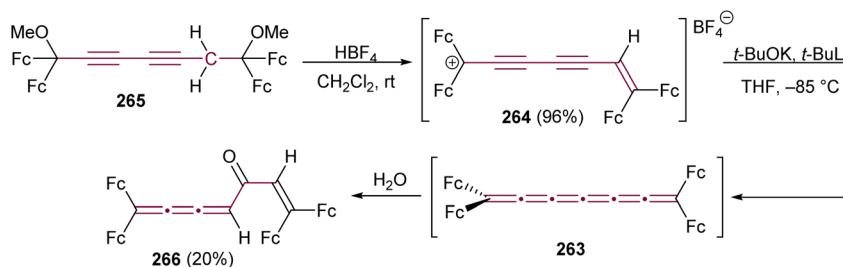


Scheme 57 Iodination and mercuriation of [4]cumulene **227**.²¹⁸



Table 1 Relative π -electron energies of tetraphenyl[n]cumulenes and their conjugate acids²¹⁸

Cumulene/conjugate acid	Total π -electron energy	π -Electron energy loss resulting from protonation
$\text{Ph}_2\text{C}=\text{C}=\text{CPh}_2/\text{Ph}_2\text{C}=\text{CH}^+\text{CPh}_2$	36.88/35.90 β	0.98 β
$\text{Ph}_2\text{C}=\text{C}=\text{C}=\text{CPh}_2/\text{Ph}_2\text{C}=\text{C}=\text{CH}^+\text{CPh}_2$	39.38/37.82 β	1.56 β
$\text{Ph}_2\text{C}=\text{C}=\text{C}=\text{C}=\text{CPh}_2/\text{Ph}_2\text{C}=\text{C}=\text{C}=\text{CH}^+\text{CPh}_2$	41.88/40.80 β	1.08 β
$\text{Ph}_2\text{C}=\text{C}=\text{C}=\text{C}=\text{C}=\text{CPh}_2/\text{Ph}_2\text{C}=\text{C}=\text{C}=\text{C}=\text{CH}^+\text{CPh}_2$	44.40/42.90 β	1.50 β

Scheme 58 *In situ* formation of [6]cumulene **263** followed by hydrolysis to give **266** (Fc = ferrocenyl).²¹⁷

could not be characterized, although it is noted that three molar equivalents of Br_2 are consumed during the bromination.

Similar to reactions of odd [n]cumulenes with HgCl_2 , treating tetraphenyl[4]cumulene **200** with HgCl_2 reportedly fails to produce stable products that could be either isolated or studied.²¹⁸ In contrast, the reaction of **227** with HgCl_2 immediately yields a purple crystalline precipitate in 70% yield that is identified as complex **262** formed *via* addition to the β -bond (Scheme 57). Spectroscopic analyses confirm that **262** is structurally similar to **261**, exhibiting essentially identical IR and UV-vis absorption characteristics. The formation of **262** is reversible, and shaking **262** in a water or alcohol solution regenerates **227**, which crystallizes out of the solution.²¹⁸

The combined studies by Fischer and Fischer reveal key reactivity trends, particularly regarding protonation reactions.²¹⁸ Most notably, allenes and even [n]cumulenes ($n = 4$) exhibit enhanced reactivity compared to odd [n]cumulenes ($n = 3$ and 5). Based on the evidence in hand, it also seems clear that the β -bond in [4]cumulenes is more reactive than the terminal α -bond. Fischer and Fischer rationalize this trend based on the loss of π -electron energy upon, *e.g.*, protonation. This analysis clearly shows that the energy loss is more significant for even than odd [n]cumulenes (Table 1). While not reported in the study, this interpretation could easily be extended to other electrophilic reactivity patterns that show preference for position C2. Thus, the experimental outcomes align with the hypothesis that the enhanced reactivity of the β -bond stems from its lower reorganization energy, leading to a preference for this bond in the reaction pathway and ultimately producing a stable product.

Although the isolation and characterization of a [6]cumulene have not yet been reported, a single study by Bildstein and coworkers from 2001 describes the *in situ* formation of tetraferrocenyl[6]cumulene **263** (Scheme 58).²¹⁷ Their efforts to establish the formation of **263** include the elimination reaction

of diene **265** to form the cation **264**, followed by a trapping experiment with H_2O , giving **266** as the product.

Conclusions

Over the past century, the chemistry of [n]cumulenes has witnessed the development of a variety of synthetic methodologies that allow the construction of these “exotic” molecules. Historically, much of the motivation has stemmed from studying [n]cumulenes as models of carbyne-like systems or from harnessing their unique electronic properties. However, the robust synthetic protocols now available, particularly for [3]- and [5]cumulenes, highlight that these systems can also serve as remarkably versatile building blocks in organic synthesis.

Despite the challenges that arise when attempting to access longer odd (beyond [5]cumulenes) and even [n]cumulenes, the chemistry surveyed in this review demonstrates that these compounds are far from being mere curiosities. Indeed, the high internal energy inherent in [n]cumulenes leads to reactivity profiles that can unlock new transformations and provide innovative routes to frameworks otherwise difficult to obtain. From cyclooligomerization and cycloaddition processes to organometallic transformations, [n]cumulenes repeatedly display a capacity to form carbon skeletons that are largely inaccessible from more conventional substrates.

We hope that this comprehensive overview will convince the broader organic chemistry community to consider [n]cumulenes as useful synthetic building blocks. By incorporating cumulene moieties into retrosynthetic plans, chemists can exploit their unusual reactivity to generate novel architectures and potentially uncover uncharted reaction pathways. In this way, [n]cumulenes offer a potent combination of structural diversity, synthetic accessibility (especially for the shorter members), and distinctive chemical behavior, an attractive toolkit that deserves wider recognition and application in contemporary organic synthesis.



Author contributions

All authors contributed to writing and editing this article.

Conflicts of interest

The authors declare no competing interests.

Data availability

No primary research results, software or code have been included and no new data were generated or analysed as part of this review.

Acknowledgements

P. G. and A. P. acknowledge funding from the Polish National Science Centre (SONATA 2020/39/D/ST4/00560) and the National Agency for Academic Exchange (Polish Returns PPN_PPO_2020_1_00012). R. R. T. acknowledges funding from the Natural Sciences and Engineering Research Council of Canada (NSERC RGPIN-2017-05052 and RGPIN-2023-04000).

References

- 1 A. D. McNaught, A. Wilkinson, I. U. o. Pure and A. Chemistry, *Compendium of Chemical Terminology: IUPAC Recommendations*, Blackwell Scientific Publications, Oxford, 1997.
- 2 J. H. van't Hoff, *La chimie dans l'espace*, Bazendijk, Rotterdam, 1875.
- 3 J. H. van't Hoff, *The arrangement of atoms in space*, Longmans, Green and Co., London, 1898.
- 4 F. Diederich and M. Kivala, *Adv. Mat.*, 2010, **22**, 803–812.
- 5 M. J. Rice, S. R. Phillpot, A. R. Bishop and D. K. Campbell, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1986, **34**, 4139–4149.
- 6 M. Liu, V. I. Artyukhov, H. Lee, F. Xu and B. I. Yakobson, *ACS Nano*, 2013, **7**, 10075–10082.
- 7 V. I. Artyukhov, M. Liu and B. I. Yakobson, *Nano Lett.*, 2014, **14**, 4224–4229.
- 8 K. Brand, *Ber. Dtsch. Chem. Ges.*, 1921, **54**, 1987–2006.
- 9 K. Brand, F. Kercher and B. Dtsch, *Chem. Ges.*, 1921, **54**, 2007–2017.
- 10 R. Hoffmann, *Tetrahedron*, 1966, **22**, 521–538.
- 11 S. Yang and M. Kertesz, *J. Phys. Chem. A*, 2007, **112**, 146–151.
- 12 C. Neiss, E. Trushin and A. Görling, *ChemPhysChem*, 2014, **15**, 2497–2502.
- 13 H. L. Anderson, C. W. Patrick, L. M. Scriven and S. L. Woltering, *Bull. Chem. Soc. Jpn.*, 2020, **94**, 798–811.
- 14 M. Franz, J. A. Januszewski, D. Wendinger, C. Neiss, L. D. Movsisyan, F. Hampel, H. L. Anderson, A. Görling and R. R. Tykwinski, *Angew. Chem., Int. Ed.*, 2015, **54**, 6645–6649.
- 15 J. A. Januszewski, D. Wendinger, C. D. Methfessel, F. Hampel and R. R. Tykwinski, *Angew. Chem., Int. Ed.*, 2013, **52**, 1817–1821.
- 16 W. A. Chalifoux and R. R. Tykwinski, *Nat. Chem.*, 2010, **2**, 967–971.
- 17 Y. Gao, Y. Hou, F. Gordillo Gámez, M. J. Ferguson, J. Casado and R. R. Tykwinski, *Nat. Chem.*, 2020, **12**, 1143–1149.
- 18 C. W. Patrick, Y. Gao, P. Gupta, A. L. Thompson, A. W. Parker and H. L. Anderson, *Nat. Chem.*, 2024, **16**, 193–200.
- 19 K. Kaiser, L. M. Scriven, F. Schulz, P. Gawel, L. Gross and H. L. Anderson, *Science*, 2019, **365**, 1299–1301.
- 20 L. M. Scriven, K. Kaiser, F. Schulz, A. J. Sterling, S. L. Woltering, P. Gawel, K. E. Christensen, H. L. Anderson and L. Gross, *J. Am. Chem. Soc.*, 2020, **142**, 12921–12924.
- 21 Y. Gao, F. Albrecht, I. Rončević, I. Etedgui, P. Kumar, L. M. Scriven, K. E. Christensen, S. Mishra, L. Righetti, M. Rossmannek, I. Tavernelli, H. L. Anderson and L. Gross, *Nature*, 2023, **623**, 977–981.
- 22 L. Sun, W. Zheng, W. Gao, F. Kang, M. Zhao and W. Xu, *Nature*, 2023, **623**, 972–976.
- 23 P. Gawel and C. Foroutan-Nejad, *Nature*, 2023, **623**, 922–924.
- 24 S. Housmansl and G. H. P. Honnef, *Nachr. Chem., Tech. Lab.*, 1984, **32**, 379–381.
- 25 L. Leroyer, V. Maraval and R. Chauvin, *Chem. Rev.*, 2012, **112**, 1310–1343.
- 26 I. V. Alabugin and E. Gonzalez-Rodriguez, *Acc. Chem. Res.*, 2018, **51**, 1206–1219.
- 27 B. Pigulski, N. Gulia and S. Szafert, *Eur. J. Org. Chem.*, 2019, **2019**, 1420–1445.
- 28 A. D. Senese and W. A. Chalifoux, *Molecules*, 2019, **24**, 118.
- 29 P. Gawel, Y.-L. Wu, A. D. Finke, N. Trapp, M. Zalibera, C. Boudon, J.-P. Gisselbrecht, W. B. Schweizer, G. Gescheidt and F. Diederich, *Chem. Eur. J.*, 2015, **21**, 6215–6225.
- 30 Y.-L. Wu, F. Bureš, P. D. Jarowski, W. B. Schweizer, C. Boudon, J.-P. Gisselbrecht and F. Diederich, *Chem. Eur. J.*, 2010, **16**, 9592–9605.
- 31 R. Chauvin, C. Lepetit, V. Maraval and L. Leroyer, *Pure Appl. Chem.*, 2010, **82**, 769–800.
- 32 Z.-i. Yoshida, *Pure Appl. Chem.*, 1982, **54**, 1059–1074.
- 33 D. Bouvy, Z. Janousek, H. G. Viehe, B. Tinant and J.-P. Declercq, *Tetrahedron Lett.*, 1993, **34**, 1779–1782.
- 34 B. Tinant, J.-P. Declercq, D. Bouvy, Z. Janousek and H. G. Viehe, *J. Chem. Soc., Perkin Trans. 2*, 1993, 911–915.
- 35 W. M. Schubert, T. H. Liddicoet and W. A. Lanka, *J. Am. Chem. Soc.*, 1952, **74**, 569.
- 36 P. A. Morken, P. C. Bachand, D. C. Swenson and D. J. Burton, *J. Am. Chem. Soc.*, 1993, **115**, 5430–5439.
- 37 A. Roedig and G. Zaby, *Chem. Ber.*, 1980, **113**, 3342–3351.
- 38 J. Woliński, *Rocz. Chem.*, 1957, **31**, 1189–1198.
- 39 S. Kishigami, K. Tanaka and F. Toda, *Chem. Lett.*, 1990, **19**, 1877–1880.
- 40 P. A. Morken, D. J. Burton and D. C. Swenson, *J. Org. Chem.*, 1994, **59**, 2119–2121.
- 41 D. Wu, Y. Li, R. Ganguly and R. Kinjo, *Chem. Commun.*, 2014, **50**, 12378–12381.
- 42 R. H. Baughman, *Science*, 2006, **312**, 1009–1010.



- 43 R. O. Uhler, H. Shechter and G. V. D. Tiers, *J. Am. Chem. Soc.*, 1962, **84**, 3397–3398.
- 44 Z. Berkovitch-Yellin, M. Lahav and L. Leiserowitz, *J. Am. Chem. Soc.*, 1974, **96**, 918–920.
- 45 B. Heinrich and A. Roedig, *Angew. Chem. Int. Ed. Engl.*, 1968, **7**, 375–376.
- 46 S. Koster and R. West, *J. Chem. Soc. D*, 1971, 1380.
- 47 S. K. Koster and R. West, *J. Org. Chem.*, 1975, **40**, 2300–2304.
- 48 F. W. Nader, C.-D. Wacker, H. Irngartinger, U. Huber-Patz, R. Jahn and H. Rodewald, *Angew. Chem., Int. Ed. Engl.*, 1985, **24**, 852–853.
- 49 H. Uno, K.-i. Kasahara, N. Nibu, S.-i. Nagaoka and N. Ono, *J. Org. Chem.*, 2000, **65**, 1615–1622.
- 50 L. Hagelee, R. West, J. Calabrese and J. Norman, *J. Am. Chem. Soc.*, 1979, **101**, 4888–4892.
- 51 M. Iyoda, S. Tanaka, M. Nose and M. Oda, *Chem. Commun.*, 1983, 1058–1059.
- 52 M. Iyoda, S. Tanaka, H. Otani, M. Nose and M. Oda, *J. Am. Chem. Soc.*, 1988, **110**, 8494–8500.
- 53 P. J. Stang, M. R. White and G. Maas, *Organometallics*, 1983, **2**, 720–725.
- 54 L. Stehling and G. Wilke, *Angew. Chem., Int. Ed. Engl.*, 1985, **24**, 496–497.
- 55 J.-D. van Loon, P. Seiler and F. Diederich, *Angew. Chem., Int. Ed. Engl.*, 1993, **32**, 1187–1189.
- 56 Y.-L. Wu, F. Tancini, W. B. Schweizer, D. Paunescu, C. Boudon, J.-P. Gisselbrecht, P. D. Jarowski, E. Dalcanale and F. Diederich, *Chem.-Asian J.*, 2012, **7**, 1185–1190.
- 57 H. D. Hartzler, *J. Am. Chem. Soc.*, 1966, **88**, 3155–3156.
- 58 H. D. Hartzler, *J. Am. Chem. Soc.*, 1971, **93**, 4527–4531.
- 59 L. T. Scott and G. J. DeCicco, *Tetrahedron Lett.*, 1976, **17**, 2663–2666.
- 60 C. Santiago, K. N. Houk, G. J. DeCicco and L. T. Scott, *J. Am. Chem. Soc.*, 1978, **100**, 692–696.
- 61 M. Kaftory, I. Agmon, M. Ladika and P. J. Stang, *J. Am. Chem. Soc.*, 1987, **109**, 782–787.
- 62 W. M. Schubert, T. H. Liddicoet and W. A. Lanka, *J. Am. Chem. Soc.*, 1954, **76**, 1929–1932.
- 63 E. Kloster-Jensen and J. Wirz, *Helv. Chim. Acta*, 1975, **58**, 162–177.
- 64 E. Kloster-Jensen and J. Wirz, *Angew. Chem., Int. Ed. Engl.*, 1973, **12**, 671.
- 65 C. Werner, H. Hopf, I. Dix, P. Bubenitschek and P. G. Jones, *Chem. Eur. J.*, 2007, **13**, 9462–9477.
- 66 P. J. Stang and M. Ladika, *J. Am. Chem. Soc.*, 1980, **102**, 5406–5407.
- 67 P. J. Stang and M. Ladika, *J. Am. Chem. Soc.*, 1981, **103**, 6437–6443.
- 68 M. Iyoda, Y. Kuwatani and M. Oda, *J. Am. Chem. Soc.*, 1989, **111**, 3761–3762.
- 69 M. Iyoda, Y. Kuwatani, M. Oda, Y. Kai, N. Kaneshisa and N. Kasai, *Angew. Chem., Int. Ed. Engl.*, 1990, **29**, 1062–1064.
- 70 M. Iyoda, M. Oda, Y. Kai, N. Kaneshisa and N. Kasai, *Chem. Lett.*, 1990, **19**, 2149–2152.
- 71 Y. Kuwatani, G. Yamamoto, M. Oda and M. Iyoda, *Bull. Chem. Soc. Jpn.*, 2005, **78**, 2188–2208.
- 72 T. Nishiuchi, S. Takeuchi and T. Kubo, *Chem. Asian J.*, 2025, e202401868.
- 73 N. Islam, T. Ooi, T. Iwasawa, M. Nishiuchi and Y. Kawamura, *Chem. Commun.*, 2009, 574–576.
- 74 T. Kawase, Y. Minami, N. Nishigaki, S. Okano, H. Kurata and M. Oda, *Angew. Chem., Int. Ed.*, 2005, **44**, 316–319.
- 75 D. Wendinger, J. A. Januszewski, F. Hampel and R. R. Tykwinski, *Chem. Commun.*, 2015, **51**, 14877–14880.
- 76 W. Ried and R. Neidhardt, *Adv. Cycloaddit.*, 1970, **739**, 155–158.
- 77 H. Nemoto and K. Fukumoto, *Tetrahedron*, 1998, **54**, 5425–5464.
- 78 J. L. Segura and N. Martín, *Chem. Rev.*, 1999, **99**, 3199–3246.
- 79 A. Shimizu, S. Nobusue, H. Miyoshi and Y. Tobe, *Pure Appl. Chem.*, 2014, **86**, 517–528.
- 80 P.-H. Liu, L. Li, J. A. Webb, Y. Zhang and N. S. Goroff, *Org. Lett.*, 2004, **6**, 2081–2083.
- 81 E. L. Martin and W. H. Sharkey, *J. Am. Chem. Soc.*, 1959, **81**, 5256–5258.
- 82 C. Ehm and D. Lentz, *Chem. Commun.*, 2010, **46**, 2399–2401.
- 83 M. I. Bruce, J. R. Rodgers, M. R. Snow and A. G. Swincer, *Chem. Commun.*, 1981, 271–272.
- 84 T. Michinobu and F. Diederich, *Angew. Chem., Int. Ed.*, 2018, **57**, 3552–3577.
- 85 F. Bureš, W. B. Schweizer, J. C. May, C. Boudon, J.-P. Gisselbrecht, M. Gross, I. Biaggio and F. Diederich, *Chem. Eur. J.*, 2007, **13**, 5378–5387.
- 86 M. Kivala, C. Boudon, J.-P. Gisselbrecht, B. Enko, P. Seiler, I. B. Müller, N. Langer, P. D. Jarowski, G. Gescheidt and F. Diederich, *Chem. Eur. J.*, 2009, **15**, 4111–4123.
- 87 Y.-L. Wu, P. D. Jarowski, W. B. Schweizer and F. Diederich, *Chem. Eur. J.*, 2010, **16**, 202–211.
- 88 B. Breiten, Y.-L. Wu, P. D. Jarowski, J.-P. Gisselbrecht, C. Boudon, M. Griesser, C. Onitsch, G. Gescheidt, W. B. Schweizer, N. Langer, C. Lennartz and F. Diederich, *Chem. Sci.*, 2011, **2**, 88–93.
- 89 M. Štefko, M. D. Tzirakis, B. Breiten, M.-O. Ebert, O. Dumele, W. B. Schweizer, J.-P. Gisselbrecht, C. Boudon, M. T. Beels, I. Biaggio and F. Diederich, *Chem. Eur. J.*, 2013, **19**, 12693–12704.
- 90 B. H. Tchitchanov, M. Chiu, M. Jordan, M. Kivala, W. B. Schweizer and F. Diederich, *Eur. J. Org. Chem.*, 2013, **2013**, 3729–3740.
- 91 T. Michinobu, J. C. May, J. H. Lim, C. Boudon, J.-P. Gisselbrecht, P. Seiler, M. Gross, I. Biaggio and F. Diederich, *Chem. Commun.*, 2005, **46**, 737–739.
- 92 S.-i. Kato and F. Diederich, *Chem. Commun.*, 2010, 1994–2006.
- 93 M. I. Bruce, *Aust. J. Chem.*, 2011, **64**, 77–103.
- 94 T. Shoji, M. Maruyama, E. Shimomura, A. Maruyama, S. Ito, T. Okujima, K. Toyota and N. Morita, *J. Org. Chem.*, 2013, **78**, 12513–12524.
- 95 M. Betou, N. Kerisit, E. Meledje, Y. R. Leroux, C. Katan, J.-F. Halet, J.-C. Guillemin and Y. Trolez, *Chem. Eur. J.*, 2014, **20**, 9553–9557.
- 96 H. Maurer and H. Hopf, *Angew. Chem. Int. Ed. Engl.*, 1976, **15**, 628–629.



- 97 H. Hopf, *Chem. Ber.*, 1971, **104**, 3087–3095.
- 98 A. Baroudi, W. E. Matar and S. Tlais, *ACS Omega*, 2024, **9**, 41884–41889.
- 99 J. E. Douglas, B. S. Rabinovitch and F. S. Looney, *J. Chem. Phys.*, 1955, **23**, 315–323.
- 100 N. Islam, M. Tsukayama and Y. Kawamura, *Int. J. Mod. Phys. B*, 2006, **20**, 4619–4624.
- 101 S. Ueta, K. Hida, M. Nishiuchi and Y. Kawamura, *Org. Biomol. Chem.*, 2014, **12**, 2784–2791.
- 102 P. Gawel, C. Dengiz, A. D. Finke, N. Trapp, C. Boudon, J.-P. Gisselbrecht and F. Diederich, *Angew. Chem., Int. Ed.*, 2014, **53**, 4341–4345.
- 103 N. Kerisit, P. Gawel, B. Levandowski, Y. F. Yang, V. García-López, N. Trapp, L. Ruhlmann, C. Boudon, K. N. Houk and F. Diederich, *Chem. Eur. J.*, 2018, **24**, 159–168.
- 104 E. A. Margulies, Y.-L. Wu, P. Gawel, S. A. Miller, L. E. Shoer, R. D. Schaller, F. Diederich and M. R. Wasielewski, *Angew. Chem., Int. Ed.*, 2015, **54**, 8679–8683.
- 105 E. A. Margulies, N. Kerisit, P. Gawel, C. M. Mauck, L. Ma, C. E. Miller, R. M. Young, N. Trapp, Y.-L. Wu, F. Diederich and M. R. Wasielewski, *J. Phys. Chem. C*, 2017, **121**, 21262–21271.
- 106 J. A. Januszewski, F. Hampel, C. Neiss, A. Görling and R. R. Tykwinski, *Angew. Chem., Int. Ed.*, 2014, **53**, 3743–3747.
- 107 B. Bildstein, M. Schweiger, H. Angleitner, H. Kopacka, K. Wurst, K.-H. Ongania, M. Fontani and P. Zanello, *Organometallics*, 1999, **18**, 4286–4295.
- 108 R. G. Visser and H. J. T. Bos, *Tetrahedron Lett.*, 1979, **20**, 4857–4858.
- 109 A. C. Brouwer, A. V. E. George, D. Seykens and H. J. T. Bos, *Tetrahedron Lett.*, 1978, **19**, 4839–4840.
- 110 K. C. Nicolaou and W. M. Dai, *Angew. Chem., Int. Ed. Engl.*, 1991, **30**, 1387–1416.
- 111 I. H. Goldberg, *Acc. Chem. Res.*, 1991, **24**, 191–198.
- 112 K. C. Nicolaou, A. L. Smith and E. W. Yue, *Proc. Natl. Acad. Sci. U. S. A.*, 1993, **90**, 5881–5888.
- 113 K. Fujiwara, H. Sakai and M. Hiram, *J. Org. Chem.*, 1991, **56**, 1688–1689.
- 114 J. Gabriel Garcia, B. Ramos, L. M. Pratt and A. Rodríguez, *Tetrahedron Lett.*, 1995, **36**, 7391–7394.
- 115 X.-Y. Guan and M. Shi, *J. Org. Chem.*, 2009, **74**, 1977–1981.
- 116 Z. Xu and X. Lu, *J. Org. Chem.*, 1998, **63**, 5031–5041.
- 117 L.-F. Wang, X.-P. Cao, Z.-F. Shi, P. An and H.-F. Chow, *Adv. Synth. Catal.*, 2014, **356**, 3383–3390.
- 118 W. Wirschun, G.-M. Maier and J. C. Jochims, *Tetrahedron*, 1997, **53**, 5755–5766.
- 119 C. C. C. Johansson Seechurn, M. O. Kitching, T. J. Colacot and V. Snieckus, *Angew. Chem., Int. Ed.*, 2012, **51**, 5062–5085.
- 120 A. Nakamura, P.-J. Kim and N. Hagihara, *J. Organomet. Chem.*, 1965, **3**, 7–15.
- 121 A. Nakamura, *Bull. Chem. Soc. Jpn.*, 1965, **38**, 1868–1873.
- 122 D. Bright and O. S. Mills, *J. Chem. Soc. A*, 1971, 1979–1982.
- 123 D. Bright and O. S. Mills, *J. Chem. Soc., Dalton Trans.*, 1972, 2465–2469.
- 124 R. B. King and C. A. Harmon, *J. Organomet. Chem.*, 1975, **88**, 93–100.
- 125 U. Schubert and J. Grönen, *Chem. Ber.*, 1989, **122**, 1237–1245.
- 126 R. A. Fischer, R. W. Fischer, W. A. Herrmann and E. Herdtweck, *Chem. Ber.*, 1989, **122**, 2035–2040.
- 127 L. Song, A. M. Arif and P. J. Stang, *J. Organomet. Chem.*, 1990, **395**, 219–226.
- 128 M. Iyoda, Y. Kuwatani, M. Oda, K. Tatsumi and A. Nakamura, *Angew. Chem., Int. Ed. Engl.*, 1991, **30**, 1670–1672.
- 129 M. Iyoda, Y. Kuwatani and M. Oda, *J. Chem. Soc., Chem. Commun.*, 1992, 399–400.
- 130 H. Werner, M. Laubender, R. Wiedemann and B. Windmüller, *Angew. Chem., Int. Ed. Engl.*, 1996, **35**, 1237–1239.
- 131 H. Werner, R. Wiedemann, N. Mahr, P. Steinert and J. Wolf, *Chem. Eur. J.*, 1996, **2**, 561–569.
- 132 I. Kovacic, M. Laubender and H. Werner, *Organometallics*, 1997, **16**, 5607–5609.
- 133 A. Zimniak and G. Bakalarski, *J. Organomet. Chem.*, 2001, **634**, 198–208.
- 134 H. Werner, R. Wiedemann, M. Laubender, B. Windmüller, P. Steinert, O. Gevert and J. Wolf, *J. Am. Chem. Soc.*, 2002, **124**, 6966–6980.
- 135 N. Suzuki, M. Nishiura and Y. Wakatsuki, *Science*, 2002, **295**, 660–663.
- 136 R. P. Hughes, R. B. Laritchev, L. N. Zakharov and A. L. Rheingold, *J. Am. Chem. Soc.*, 2004, **126**, 2308–2309.
- 137 N. Suzuki, T. Watanabe, T. Hirose and T. Chihara, *J. Organomet. Chem.*, 2007, **692**, 5317–5321.
- 138 N. Suzuki, *J. Synth. Org. Chem., Jpn.*, 2007, **65**, 347–357.
- 139 N. Suzuki and B. Jap, *Soc. Coord. Chem.*, 2009, **53**, 43–57.
- 140 V. Cadierno and J. Gimeno, *Chem. Rev.*, 2009, **109**, 3512–3560.
- 141 C. Coletti, A. Marrone and N. Re, *Acc. Chem. Res.*, 2012, **45**, 139–149.
- 142 A. Zweig and A. K. Hoffmann, *J. Am. Chem. Soc.*, 1962, **84**, 3278–3284.
- 143 J. M. Edinger and A. R. Day, *J. Org. Chem.*, 1971, **36**, 240–242.
- 144 R. Nahon and A. R. Day, *J. Org. Chem.*, 1965, **30**, 1973–1976.
- 145 W. Kemula and J. Kornacki, *Rocz. Chem.*, 1962, **36**, 1857–1862.
- 146 W. Kemula and J. Kornacki, *Rocz. Chem.*, 1962, **36**, 1849–1855.
- 147 W. Kemula and J. Kornacki, *Rocz. Chem.*, 1962, **36**, 1835–1847.
- 148 H. Bock, C. Arad, C. Näther, I. Göbel, A. John and R. Baur, *Z. Naturforsch. B*, 1996, **51**, 1381–1390.
- 149 S. N. Spisak, M. U. Bühringer, Z. Wei, Z. Zhou, R. R. Tykwinski and M. A. Petrukhina, *Angew. Chem., Int. Ed.*, 2019, **58**, 2023–2028.
- 150 Z. Zhou, M. A. Johnson, Z. Wei, M. U. Bühringer, M. H. Garner, R. Tykwinski and M. A. Petrukhina, *Chem. Eur. J.*, 2024, **30**, e202304145.



- 151 T. Kusakawa, Y. Kabe, B. Nestler and W. Ando, *Organometallics*, 1995, **14**, 2556–2564.
- 152 D. Seyferth, E. W. Goldman and J. Escudié, *J. Organomet. Chem.*, 1984, **271**, 337–352.
- 153 A. Maercker, H. Wunderlich and U. Girreser, *Eur. J. Org. Chem.*, 1998, **1998**, 793–798.
- 154 C. B. Ziegler, Jr., *J. Org. Chem.*, 1990, **55**, 2983–2986.
- 155 G. Dyker, S. Borowski, G. Henkel, A. Kellner, I. Dix and P. G. Jones, *Tetrahedron Lett.*, 2000, **41**, 8259–8262.
- 156 T. Asakawa, M. Iinuma, Y. Wakasugi, M. Kuno, T. Furuta, S. Fujii, K. Tanaka and T. Kan, *Heterocycles*, 2010, **80**, 1125–1147.
- 157 T. Furuta, T. Asakawa, M. Iinuma, S. Fujii, K. Tanaka and T. Kan, *Chem. Commun.*, 2006, **34**, 3648–3650.
- 158 T. Asakawa, M. Iinuma, T. Furuta, S. Fujii, T. Kan and K. Tanaka, *Chem. Lett.*, 2006, **35**, 512–513.
- 159 A. Radadiya and A. Shah, *Eur. J. Med. Chem.*, 2015, **97**, 356–376.
- 160 B. Alcaide, P. Almendros, S. Cembellín, I. Fernández and T. Martínez del Campo, *Chem. Eur. J.*, 2016, **22**, 11667–11676.
- 161 E. Wang, X. Fu, X. Xie, J. Chen, H. Gao and Y. Liu, *Tetrahedron Lett.*, 2011, **52**, 1968–1972.
- 162 Z.-Y. Xue, X. Fang and C.-J. Wang, *Org. Biomol. Chem.*, 2011, **9**, 3622–3624.
- 163 R. Ballini, A. Palmieri and M. Petrini, *Org. Chem. Front.*, 2022, **9**, 6077–6113.
- 164 Z. Liu, P. Gu and M. Shi, *Chem. Eur. J.*, 2011, **17**, 5796–5799.
- 165 A. Konishi, S. Satake and M. Yasuda, *Chem. Lett.*, 2020, **49**, 589–592.
- 166 A. Konishi, S. Imai, S. Satake, K. Chiba and M. Yasuda, *Chem. Eur. J.*, 2023, **29**, e202301255.
- 167 J. S. Rogers, R. J. Lachicotte and G. C. Bazan, *Organometallics*, 1999, **18**, 3976–3980.
- 168 F. Klotter and A. Studer, *Angew. Chem., Int. Ed.*, 2014, **53**, 2473–2476.
- 169 T. Nakano, K. Takewaki, T. Yade and Y. Okamoto, *J. Am. Chem. Soc.*, 2001, **123**, 9182–9183.
- 170 K. Brand, *Chem. Ber.*, 1950, **83**, 346–354.
- 171 N. R. Browne, R. F. C. Brown, F. W. Eastwood and G. D. Fallon, *Aust. J. Chem.*, 1987, **40**, 1675–1686.
- 172 W. Li and M. Shi, *Org. Biomol. Chem.*, 2009, **7**, 1775–1777.
- 173 P. Gawel, E. A. Halabi, D. Schweinfurth, N. Trapp, L. Ruhlmann, C. Boudon and F. Diederich, *Eur. J. Org. Chem.*, 2016, **2016**, 2919–2924.
- 174 F. Yagishita, K. Hoshi, Y. Yoshida, S. Ueta, K. Minagawa, Y. Imada and Y. Kawamura, *Eur. J. Org. Chem.*, 2021, **2021**, 235–238.
- 175 F. Yagishita, S. Mukai, S. Abe, S. Ueta, Y. Yoshida, Y. Arakawa, K. Minagawa and Y. Imada, *RSC Mechanochem.*, 2024, **1**, 318–321.
- 176 K. Hoshi, M. Yasuda, T. Nakamura, Y. Yoshida, S. Ueta, K. Minagawa, Y. Kawamura, Y. Imada and F. Yagishita, *Org. Biomol. Chem.*, 2021, **19**, 7594–7597.
- 177 S. Yang and M. Shi, *Acc. Chem. Res.*, 2018, **51**, 1667–1680.
- 178 N. Tokitoh, H. Hayakawa, M. Goto and W. Ando, *Tetrahedron Lett.*, 1988, **29**, 1935–1938.
- 179 W. Ando, N. Tokitoh and Y. Kabe, *Phosphorus Sulfur Silicon Relat. Elem.*, 1991, **58**, 179–205.
- 180 N. Tokitoh, S. Imai, M. Goto and W. Ando, *Tetrahedron Lett.*, 1989, **30**, 4825–4828.
- 181 N. Tokitoh, M. Noguchi, Y. Kabe, W. Ando, M. Goto and H. Maki, *Tetrahedron Lett.*, 1990, **31**, 7641–7644.
- 182 N. Tokitoh, T. Suzuki and W. Ando, *Tetrahedron Lett.*, 1989, **30**, 4271–4274.
- 183 T. H. Chan and B. S. Ong, *Tetrahedron*, 1980, **36**, 2269–2289.
- 184 C. S. Adams, C. D. Weatherly, E. G. Burke and J. M. Schomaker, *Chem. Soc. Rev.*, 2014, **43**, 3136–3163.
- 185 J. K. Crandall, G. E. Salazar and R. J. Watkins, *J. Am. Chem. Soc.*, 1987, **109**, 4338–4341.
- 186 T. Greibrokk and L. Skattebøl, *Acta Chem. Scand.*, 1973, **27**, 1421–1426.
- 187 J. K. Crandall, D. M. Coppert, T. Schuster and F. Lin, *J. Am. Chem. Soc.*, 1992, **114**, 5998–6002.
- 188 G. Köbrich, H. Heinemann and W. Zündorf, *Tetrahedron*, 1967, **23**, 565–584.
- 189 L. Skattebøl, *Tetrahedron Lett.*, 1965, **6**, 2175–2179.
- 190 J. C. Jochims and G. Karich, *Tetrahedron Lett.*, 1974, **15**, 4215–4218.
- 191 R. P. Johnson, *Chem. Rev.*, 1989, **89**, 1111–1124.
- 192 W. R. Moore and T. M. Ozretich, *Tetrahedron Lett.*, 1967, **8**, 3205–3207.
- 193 R. C. Burrell, K. J. Daoust, A. Z. Bradley, K. J. DiRico and R. P. Johnson, *J. Am. Chem. Soc.*, 1996, **118**, 4218–4219.
- 194 K. Saito, M. Sakura, S. Ando and A. Hattori, *Heterocycles*, 1999, **51**, 547–556.
- 195 E. R. Hickey and L. A. Paquette, *Tetrahedron Lett.*, 1994, **35**, 2309–2312.
- 196 Z.-Y. Wang, Y.-Z. Xu, Z. Cui, L. Zhang, J.-Y. Lu, Y.-Q. Zheng and R. Zhu, *Nature Synth.*, 2025, **4**, 702–709.
- 197 A. V. Kelleghan, A. S. Bulger, D. C. Witkowski and N. K. Garg, *Nature*, 2023, **618**, 748–754.
- 198 Z. G. Walters, D. C. Witkowski, K. N. Houk and N. K. Garg, *Chem. Sci.*, 2025, **16**, 16047–16056.
- 199 H. H. Wenk, M. Winkler and W. Sander, *Angew. Chem., Int. Ed.*, 2003, **42**, 502–528.
- 200 N. Kim, M. Choi, S.-E. Suh and D. M. Chenoweth, *Chem. Rev.*, 2024, **124**, 11435–11522.
- 201 N. Pavliček, B. Schuler, S. Collazos, N. Moll, D. Pérez, E. Guitián, G. Meyer, D. Peña and L. Gross, *Nat. Chem.*, 2015, **7**, 623–628.
- 202 R. Warmuth, *Angew. Chem. Int. Ed. Engl.*, 1997, **36**, 1347–1350.
- 203 Z.-Y. Wang and R. Zhu, *J. Am. Chem. Soc.*, 2023, **145**, 23755–23763.
- 204 H.-L. Sun, B. Wu, D.-Q. Liu, Z.-D. Yu, J.-J. Wang, Q. Liu, X. Liu, D. Niu, J.-H. Dou and R. Zhu, *J. Am. Chem. Soc.*, 2022, **144**, 8807–8817.
- 205 B. Wu, H.-Z. Su, Z.-Y. Wang, Z.-D. Yu, H.-L. Sun, F. Yang, J.-H. Dou and R. Zhu, *J. Am. Chem. Soc.*, 2022, **144**, 4315–4320.
- 206 Z.-L. Wang and R. Zhu, *J. Am. Chem. Soc.*, 2024, **146**, 19377–19385.



- 207 J.-J. Wang, X.-Y. Zhang, J.-Y. Bai, Q.-Y. Liu, J. Wen and R. Zhu, *Chem*, 2025, **11**, 102643.
- 208 H.-Z. Su, B. Wu, J. Wang and R. Zhu, *Giant*, 2023, **13**, 100139.
- 209 U. Rosenthal, *Chem. Eur. J.*, 2021, **27**, 17751–17760.
- 210 S. Roy, U. Rosenthal and E. D. Jemmis, *Acc. Chem. Res.*, 2014, **47**, 2917–2930.
- 211 M. A. Johnson, J. A. Meckes, M. U. Bühringer, Z. Zhou, Y. Kuwatani, M. J. Ferguson, Z. Wei, M. Iyoda, M. A. Petrukhina and R. R. Tykwinski, *J. Phys. Org. Chem.*, 2023, **36**, e4454.
- 212 D. R. Taylor, *Chem. Rev.*, 1967, **67**, 317–359.
- 213 S. Ma, *Acc. Chem. Res.*, 2009, **42**, 1679–1688.
- 214 N. Krause and C. Winter, *Chem. Rev.*, 2011, **111**, 1994–2009.
- 215 S. Yu and S. Ma, *Angew. Chem., Int. Ed.*, 2012, **51**, 3074–3112.
- 216 D. Wendinger and R. R. Tykwinski, *Acc. Chem. Res.*, 2017, **50**, 1468–1479.
- 217 B. Bildstein, W. Skibar, M. Schweiger, H. Kopacka and K. Wurst, *J. Organomet. Chem.*, 2001, **622**, 135–142.
- 218 H. Fischer and H. Fischer, *Chem. Ber.*, 1964, **97**, 2959–3264.
- 219 R. Kuhn, H. Fischer and H. Fischer, *Chem. Ber.*, 1964, **97**, 1760–1766.
- 220 K. W. Ratts and R. D. Partos, *J. Am. Chem. Soc.*, 1969, **91**, 6112–6115.
- 221 M. Iyoda, in *Science of Synthesis, Category 6, Compounds with All-Carbon Functions*, eds J. S. Siegel, Y. Tobe and I. Shinkai, Georg Thieme Verlag KG, Stuttgart, 1st edn, 2009, vol. 45a, pp. 507–545.
- 222 M. Kaftory, I. Agmon, M. Ladika and P. J. Stang, *J. Am. Chem. Soc.*, 1987, **109**, 782–787.
- 223 P. J. Stang, *Chem. Rev.*, 1978, **78**, 383–405.
- 224 W. J. le Noble, S. Basak and S. Srivastava, *J. Am. Chem. Soc.*, 1981, **103**, 4638–4639.
- 225 S. Basak, S. Srivastava and W. J. le Noble, *J. Org. Chem.*, 1987, **52**, 5095–5099.
- 226 P. J. Stang and A. E. Learned, *J. Chem. Soc., Chem. Commun.*, 1988, 301–302.
- 227 A. E. Learned, A. M. Arif and P. J. Stang, *J. Org. Chem.*, 1988, **53**, 3122–3123.
- 228 R. G. Visser, E. A. Oostveen and H. J. T. Bos, *Tetrahedron Lett.*, 1981, **22**, 1139–1140.
- 229 G. Karich and J. C. Jochims, *Chem. Ber.*, 1977, **110**, 2680–2694.
- 230 F. W. Nader and A. Brecht, *Angew. Chem. Int. Ed. Engl.*, 1986, **25**, 93–94.
- 231 M. v. Herrath and D. Rewicki, *Liebigs Ann. Chem.*, 1977, **4**, 582–587.
- 232 H. Lorenz and R. Wizinger, *Helv. Chim. Acta*, 1945, **28**, 600–612.
- 233 R. Wizinger and G. Renckhoff, *Helv. Chim. Acta*, 1941, **24**, 369–388.
- 234 Z. Zhou, M. A. Johnson, Z. Wei, M. U. Bühringer, M. H. Garner, R. R. Tykwinski and M. A. Petrukhina, *Chem. Eur. J.*, 2024, **30**, e202304145.
- 235 M. Nakagawa, K. Shingū and K. Naemura, *Tetrahedron Lett.*, 1961, **2**, 802–806.
- 236 R. Kuhn and B. Schulz, *Angew. Chem., Int. Ed.*, 1962, **1**, 211.
- 237 R. Kuhn and B. Schulz, *Chem. Ber.*, 1963, **96**, 3200–3208.
- 238 R. Kuhn and H. Fischer, *Angew. Chem., Int. Ed.*, 1963, **2**, 692–693.
- 239 K. Bertsch, M. A. Rahman and J. C. Jochims, *Chem. Ber.*, 1979, **112**, 567–576.
- 240 J.-J. Dai, X. Yin, L. Li, M. E. Rivera, Y.-C. Wang and M. Dai, *Nat. Commun.*, 2023, **14**, 1–10.

