

Cite this: *Chem. Sci.*, 2025, 16, 23357

All publication charges for this article have been paid for by the Royal Society of Chemistry

Received 11th June 2025
Accepted 22nd October 2025

DOI: 10.1039/d5sc04275e

rsc.li/chemical-science

Temperature-dependent reaction rates of quinone-alkene cycloaddition reveal that only entropy determines the rate of SPOCQ reactions

Johannes A. M. Damen,^a Jorge Escorihuela,^b Judith Firet,^a Han Zuilhof^{*ac} and Bauke Albada^{*a}

Second-order rate constants and thermodynamic activation parameters for the strain-promoted oxidation-controlled quinone (SPOCQ) click reaction of an *ortho*-quinone with various *trans*-cyclooctene and cyclooctyne reagents were determined by stopped-flow spectroscopic analysis. We substantiate the origin of the enhancements of the reaction rates in various sTCO derivatives as compared to TCO, and demonstrated that *ortho*-quinone-cycloalkene cycloadditions are fully entropy-controlled. The *endo*/*exo* differences of BCN in SPOCQ and SPAAC were also (re)evaluated, revealing absence of a difference in reactivity between these two isomers for both click reactions. Full crystallographic descriptions of *endo*-BCN-OH and DBCO combined with high-end DFT ring-strain computations confirm that entropy controls this reaction for both cycloalkenes and cycloalkynes alike.

Introduction

Strain-promoted oxidation-controlled *ortho*-quinone (SPOCQ) alkene/alkyne cycloadditions classify as a [4 + 2] inverse electron-demand Diels-Alder (IEDDA) reaction between the electron-poor diene moiety of a quinone and a strained unsaturated carbon-carbon bond of a cycloalkene or cycloalkyne. Prominent features of this click reaction are reaction rates that are orders of magnitude higher than that of the more commonly applied strain-promoted alkyne-azide cycloaddition (SPAAC), its applicability in various settings, and its inducible nature as the *ortho*-quinone can be generated from a corresponding phenol by oxidation.^{1,2} In the context of protein modification, a tyrosine residue can be oxidized enzymatically, resulting in a 'biogenic' click reaction that has been widely used to attach foreign properly functionalized moieties to a protein of interest,^{3,4} such as the conjugation of toxic agents, proteins, and oligonucleic acids to therapeutic antibodies (Fig. 1a).⁵⁻⁹ In addition, the SPOCQ reaction has been employed for hydrogel formation and for rapid and quantitative surface modification (Fig. 1b and c).¹⁰⁻¹² Despite its wide applicability and rapidly increasing use, understanding underlying thermodynamic parameters of various strained cyclic reagents in SPOCQ

has been limited to computational studies.¹³⁻¹⁶ Theoretical mechanistic studies postulated the involvement of secondary orbital interactions (SOIs) to rationalize the higher reactivity of strained alkynes ($k_2 \sim 1 \times 10^3 \text{ M}^{-1} \text{ s}^{-1}$)¹⁷ over that of the strained *trans*-cyclooctene TCO ($k_2 \sim 1 \times 10^1 \text{ M}^{-1} \text{ s}^{-1}$).^{15,16}

Our current paper extensively maps the kinetics and thermodynamic driving forces in SPOCQ click chemistry with a range of cycloalkenes (Scheme 1a), specifically *trans*-cyclooctenol (TCO-OH) and cyclopropanated strained-*trans*-cyclooctene stereoisomers (sTCO; sometimes also referred to as cyclopropanated TCO, cpTCO).¹⁸ These latter bicyclic derivatives of TCO are structurally similar to BCN while carrying a *trans*-alkene double bond instead of the alkyne triple bond, allowing direct comparison of their reactivities and thermodynamic activation parameters. Temperature-dependent stopped-flow kinetics were used to

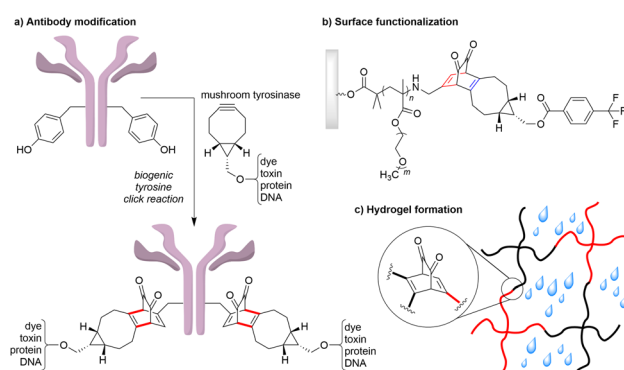
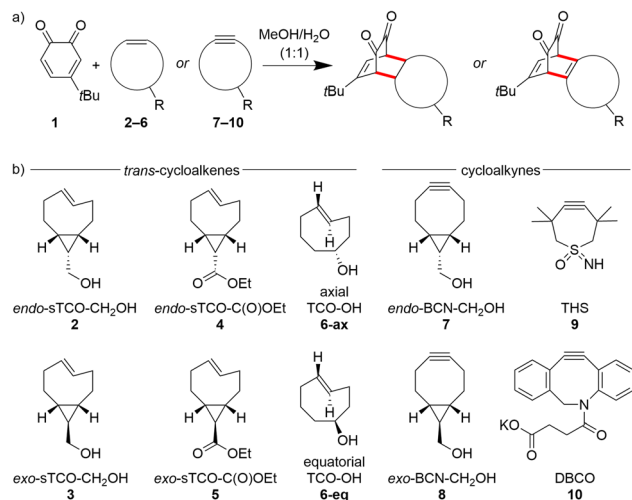


Fig. 1 Various applications of the SPOCQ click reaction. (a) in the preparation of antibody conjugates,⁵⁻⁹ (b) surface functionalization,^{11,12} and (c) the formation of polyethyleneglycol (PEG)-based hydrogels.¹⁰

^aLaboratory of Organic Chemistry, Wageningen University & Research, Stippeneng 4, Wageningen, 6807 WE, the Netherlands. E-mail: bauke.albada@wur.nl; han.zuilhof@wur.nl

^bDepartamento de Química Orgánica, Universitat de València, Facultat de Farmàcia y Ciencias de la Alimentación, Avda. Vicent Andrés Estellés s/n, València s/n, Burjassot, València, 46100, Spain

^cChina-Australia Institute for Advanced Materials and Manufacturing, Jiaying University, Jiaying 314001, China



Scheme 1 Molecules involved in the inverse electron-demand Diels–Alder reaction between an *o*-quinone diene and strained dienophiles. (a) Generalized scheme for the SPOCQ reaction with diene **1** and dienophiles containing a C=C (**2–6**) or C≡C (**7–10**) bond. (b) Matrix of the chemical structures of the cyclic dienophiles **2–10**.

acquire the second-order rate constants (k_2) whereas thermodynamic activation parameters ΔH^\ddagger and ΔS^\ddagger were obtained from Eyring analyses. With this extended matrix (Scheme 1b), the following structural features were addressed: (i) effects of chirality of exocyclic tethering points (*exo* versus *endo*); (ii) effect of axial or equatorial position of alcohol in TCO-OH; (iii) effects of different functional groups (alcohol versus ester) opposite to the dienophile; and (iv) effect of the annulation of *cis*-cyclopropane rings onto the cyclic backbone.^{18,19}

Results and discussion

Kinetic studies of the SPOCQ reaction

The second-order rate constants for the SPOCQ reaction between 4-*tert*-butyl-1,2-*ortho*-quinone **1** and dienophiles **2–10**

(Scheme 1b) were determined under pseudo-first order conditions by means of stopped-flow UV-Vis spectroscopy in a 1 : 1 (v/v) MeOH/H₂O mixtures at 25 °C. Eyring analyses were performed under similar conditions while maintaining steady temperatures at 8-centigrade intervals in the range between 5–37 °C. The obtained kinetic and thermodynamic data are provided in Table 1 and Fig. 2 (see also SI, Appendix A–D). Our ΔG^\ddagger barriers are similar to experimentally determined values for other IEDDA click reactions, *e.g.*, those between sTCO and tetrazine.¹⁸ The large negative ΔS^\ddagger values in all cases suggests that the SPOCQ cycloaddition is an associative reaction. Whereas all reactions display extremely small enthalpies of activation (ΔH^\ddagger ranges from 0.7–2.3 kcal mol^{−1}), the entropies of activation are not only larger but also vary substantially over the dataset ($T\Delta S^\ddagger$ ranges between −10.8 and −15.5 kcal mol^{−1}, see Table 1).

Eyring analysis of the SPOCQ reaction of quinone **1** with the racemic mixture of TCO-OH **6** ($k_2 = 1.2 \times 10^1 \text{ M}^{-1} \text{ s}^{-1}$) shows that the energy barrier consist of an extremely small enthalpic contribution ($\Delta H^\ddagger = 0.5 (\pm 0.1) \text{ kcal mol}^{-1}$) and a more substantial entropic contribution ($\Delta S^\ddagger = -51.9 (\pm 0.4) \text{ cal K}^{-1} \text{ mol}^{-1}$; $T\Delta S^\ddagger = -15.5 \text{ kcal mol}^{-1}$ at 25 °C), resulting in an overall Gibbs free energy of activation (ΔG^\ddagger) of 16.0 kcal mol^{−1} at 25 °C. Analysis of the two isolated diastereomers, *i.e.*, TCO-OH_{ax} (**6-ax**) and TCO-OH_{eq} (**6-eq**), revealed that the former was 12.5-times more reactive in SPOCQ with quinone **1** than the latter ($k_{2,6\text{-ax}} = 2.8 \text{ vs. } k_{2,6\text{-eq}} = 34.8 \text{ M}^{-1} \text{ s}^{-1}$), which is in line with earlier higher rates for the axial diastereomer.²⁰ Whereas Eyring analysis of the equatorial diastereomer was hampered by its slow rate, this analysis for the axial diastereomer revealed the following parameters: $\Delta H^\ddagger = 2.1 \text{ kcal mol}^{-1}$, $\Delta S^\ddagger = -44.4 \text{ cal K}^{-1} \text{ mol}^{-1}$; $T\Delta S^\ddagger = -13.2 \text{ kcal mol}^{-1}$ (at 25 °C) and an overall Gibbs free energy of activation (ΔG^\ddagger) of 15.3 kcal mol^{−1} at 25 °C.

Moving from TCO to sTCO revealed that installation of an annulated *cis*-cyclopropane ring opposite to the dienophile in *trans*-cyclooctene enhances its SPOCQ reactivity >300 fold, resulting in *exo*-sTCO-CH₂OH **3** as the most reactive reagent in

Table 1 Thermodynamic activation parameters and second-order rate constants for the inverse electron-demand Diels–Alder SPOCQ cycloaddition between 4-*tert*-butyl-1,2-*ortho*-quinone **1** and different strained cycloalkenes (**2–6**) or cycloalkynes (**7–9**) as determined by Eyring analysis in a water–MeOH mixture (1 : 1, v/v)

Dienophile	Eyring plot				k_2 plot ^b
	ΔH^\ddagger (kcal mol ^{−1})	ΔS^\ddagger (cal K ^{−1} mol ^{−1})	$T\Delta S^\ddagger,^b$ (kcal mol ^{−1})	$\Delta G^\ddagger,^b$ (kcal mol ^{−1})	k_2 (M ^{−1} s ^{−1})
Cycloalkenes					
<i>endo</i> -sTCO-CH ₂ OH (2)	0.8 (±0.2)	−40.0 (±0.7)	−11.9	12.7	(33.5 ± 0.5) · 10 ²
<i>exo</i> -sTCO-CH ₂ OH (3)	0.7 (±0.1)	−39.8 (±0.4)	−11.9	12.5	(35.3 ± 0.5) · 10 ²
<i>endo</i> -sTCO-C(O)OEt (4)	1.1 (±0.2)	−39.9 (±0.6)	−11.9	13.0	(18.7 ± 0.8) · 10 ²
<i>exo</i> -sTCO-C(O)OEt (5)	1.1 (±0.1)	−41.7 (±0.2)	−12.4	13.6	(8.2 ± 0.2) · 10 ²
TCO-OH (6)	0.5 (±0.1)	−51.9 (±0.4)	−15.5	16.0	(11.6 ± 0.1) ^a
TCO-OH axial (6-ax)	2.1 (±0.2)	−44.4 (±0.3)	−13.2	15.3	(34.8 ± 0.8)
Cycloalkynes					
<i>endo</i> -BCN-CH ₂ OH (7)	2.3 (±0.3) ^a	−36.3 (±0.9) ^a	−10.8 ^a	13.1 ^a	(18.2 ± 0.2) · 10 ^{2a}
<i>exo</i> -BCN-CH ₂ OH (8)	1.7 (±0.4)	−38.3 (±1.5)	−11.4	13.2	(16.8 ± 0.1) · 10 ²
THS (9)	0.8 (±0.2) ^a	−46.9 (±0.6) ^a	−14.0 ^a	14.8 ^a	(11.1 ± 0.2) · 10 ^{1a}

^a These values were reported previously.¹⁷ ^b Determined at 25 °C.



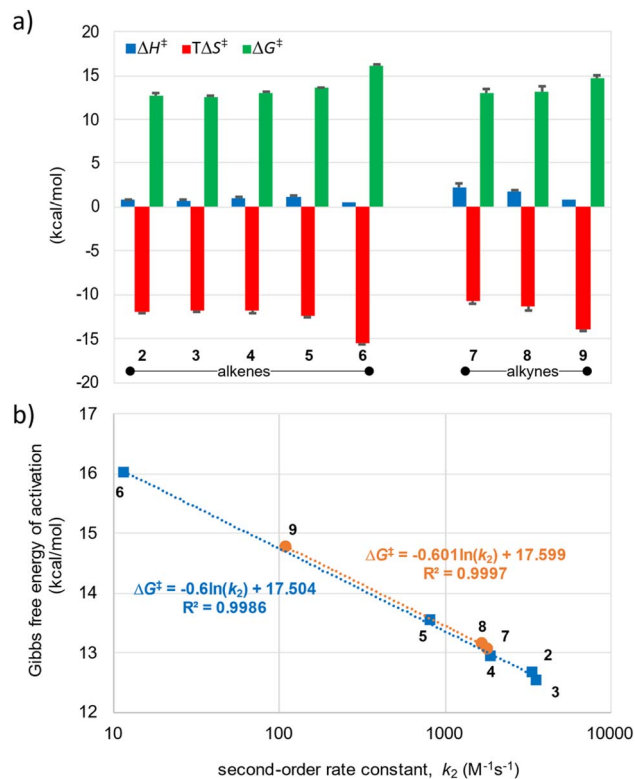


Fig. 2 (a) Thermodynamic activation parameters determined from the Eyring plots for the SPOCQ reaction between 4-*tert*-butyl-1,2-quinone (**1**) and strained alkenes (**2–6**) (left) and alkynes (**7–9**) (right). (b) High correlation of the independently determined ΔG^\ddagger and second-order rate constant for the SPOCQ reaction of the strained cycloalkenes (blue) and cycloalkynes (orange) shows accuracy of our stopped-flow method over the range of reaction rates.

the entire set. This compound displays a k_2 value of $3.5 \times 10^3 M^{-1} s^{-1}$ and thus reacts twice as fast as its *exo*-BCN-CH₂OH **8** counterpart; it is one order of magnitude slower than its reaction with a diphenyltetrazine ($k_2 = 3.3 \times 10^4 M^{-1} s^{-1}$).²¹ Interestingly, the enthalpies of activation are near-identical for TCO-OH **6** and *exo*-sTCO-CH₂OH **3** ($\Delta\Delta H^\ddagger = 0.1$ kcal mol⁻¹, which is within the standard deviation of both values, see Table 1).

Our data shows that the difference in the entropic component of the barrier accounts for the difference in free energy of activation. Specifically, the value for $\Delta(T\Delta S^\ddagger)$ for TCO-OH **6** is 3.6 kcal mol⁻¹ larger than that for *exo*-sTCO-CH₂OH **3**, resulting in $\Delta\Delta G^\ddagger = -3.4$ kcal mol⁻¹. From this we conclude that the reactant complex of *o*-quinone **1** with *exo*-sTCO-CH₂OH **3** is already more preorganized towards the transition state (TS) when compared to TCO-OH **6**. A similar analysis can be made when TCO-OH **6** is compared to *endo*-sTCO-CH₂OH **2**. As such, our results support the hypothesis by Fox that the *cis*-fused cyclopropane on sTCO results in a “half-chair”-conformation that is higher in enthalpy than the typical TCO crown conformation,¹⁸ and better preorganized towards the TS, resulting in a lower entropic barrier. For all our tested sTCO derivatives a substantially lower entropic barrier was found than for TCO-OH **6** (difference ranging from 3.1–3.6 kcal mol⁻¹), whilst

maintaining only minimal enthalpic barriers (ranging from 0.7–1.1 kcal mol⁻¹). Apparently, preorganization towards the TS enhances the rates of these dienophiles in SPOCQ.

We also measured the SPOCQ reaction to DBCO **10**, derivatives of which are frequently used in SPAAC-based reactions while sparingly being applied in SPOCQ-based bioconjugation.^{22–24} For this, pseudo-first order conditions with a 10-fold excess of DBCO-acid **10** (as potassium salt) were applied, resulting in a k_2 value of $1.7 \times 10^{-1} M^{-1} s^{-1}$ (corresponding to a calculated Gibbs free energy of activation of approximately 18.6 kcal mol⁻¹, based on the equation in Fig. 2b), which is four orders of magnitude less reactive than *endo*-BCN-CH₂OH **7**. Eyring analysis was not possible due to the substantially slower rate and associated requirement for stringent long-term temperature control.

To address the effect of the substituents opposite to the dienophile, the ester analogues of sTCO, *endo*-sTCO-C(O)OEt **4** and *exo*-sTCO-C(O)OEt **5** were measured. This revealed similar high reactivity in SPOCQ reactions, albeit lower than their alcohol counterparts. Interestingly, we observed a two-fold difference in reactivity between *endo*-sTCO-C(O)OEt **4** ($1.9 \times 10^3 M^{-1} s^{-1}$) and *exo*-sTCO-C(O)OEt **5** ($0.8 \times 10^3 M^{-1} s^{-1}$), even in repeated experiments using new batches of these dienophiles. For *exo*-sTCO-C(O)OEt **5** an additional increase in entropy of activation is found, while the enthalpy of activation is the same as for the *endo* counterpart ($\Delta H^\ddagger = 1.1$ kcal mol⁻¹). This slightly higher entropic barrier ($\Delta(T\Delta S^\ddagger) = 0.5$ kcal mol⁻¹) is reflected in the overall reaction barrier ($\Delta\Delta G^\ddagger = 0.6$ kcal mol⁻¹), resulting in the observed *ca.* 2.3-fold difference in reactivity.²⁵ The difference in activation entropy between the *endo* and *exo* isomer is tentatively attributed to interaction of the C(O)OEt group with the alkene, although the exact origin is unknown.

The effect of solvent on the kinetics and associated thermodynamic activation parameters was analyzed on a subset of dienophiles in MeOH. First, the second-order rate constant for the most hydrophobic dienophiles, *i.e.*, *endo*-sTCO-C(O)OEt **4** and *exo*-sTCO-C(O)OEt **5** (25 °C), revealed that k_2 values were an order of magnitude lower than in the water–MeOH mixture ($k_{2,(4)} = 9.5$ vs. $k_{2,(5)} = 39.6 M^{-1} s^{-1}$). Interestingly, the second-order rate constants for *endo*-sTCO-CH₂OH **2** and *exo*-sTCO-CH₂OH **3** in MeOH were in a similar range, with $k_{2,(2)} = 58.0 M^{-1} s^{-1}$ and $k_{2,(3)} = 32.9 M^{-1} s^{-1}$, but also now showing a notable difference between *endo* and *exo*. Eyring analysis of sTCO **2** and **3** revealed a slightly higher but still small enthalpic contribution for both ($\Delta H^\ddagger = 1.1$ kcal mol⁻¹) and a more pronounced differences in the larger entropic contribution ($\Delta S^\ddagger = -46.8$ cal K⁻¹ mol⁻¹ for *endo*-sTCO-CH₂OH **2** and $\Delta S^\ddagger = -47.8$ cal K⁻¹ mol⁻¹ for *exo*-sTCO-CH₂OH **3**; which corresponds to $T\Delta S^\ddagger$ values at 25 °C of -14.0 kcal mol⁻¹ for *endo*-sTCO-CH₂OH **2** and -14.3 kcal mol⁻¹ for *exo*-sTCO-CH₂OH **3**).²⁶ This result in an overall Gibbs free energy of activation (ΔG^\ddagger) of 15.1 kcal mol⁻¹ for *endo*-sTCO-CH₂OH **2** and of 15.4 kcal mol⁻¹ for *exo*-sTCO-CH₂OH **3** (at 25 °C). The increased entropic values point to a more ordered transition state in MeOH when compared to the water–MeOH mixture (1 : 1, v/v).²⁶



Revisiting the influence of BCN isomer in SPAAC

As we observed that the *exo* and *endo* isomers of sTCO-CH₂OH did not display a significant difference in their reactivity towards *o*-quinone **1** nor in the underlying activation parameters between its isomers ($k_{\text{rel},\text{exo/endo}} = 1.1$), which was previously also observed for their reactivity towards tetrazines,²¹ we evaluated the kinetics of the *exo* and *endo* isomers of the eight-membered cycloalkyne BCN. First, we determined the SPOCQ kinetics of *exo*-BCN-CH₂OH **8** with diene **1**, and compare this with our previously reported values for *endo*-BCN-CH₂OH **7**.¹⁷ We found that *exo*-BCN-CH₂OH **8** exhibits a second-order rate constant k_2 of $1.7 \times 10^3 \text{ M}^{-1} \text{ s}^{-1}$, which is near-identical to what we determined for *endo*-BCN-CH₂OH **7** (k_2 of $1.8 \times 10^3 \text{ M}^{-1} \text{ s}^{-1}$). Also, we did not detect a notable difference in the thermodynamic activation parameters associated to both isomers. This is in sharp contrast to an earlier claim regarding a notable difference in reactivity between the *exo/endo* isomers of BCN-CH₂OH in SPAAC reactions.¹⁹ Therefore, we revisited this claim by determining the reactivity of both BCN isomers with an organic azide in our stopped-flow UV-Vis equipment. As overlap of the absorption bands of both the azide and the formed triazole product at 213 nm prevented UV-Vis measurements of this transformation, we used the known fluorogenic substrate 3-azido-7-hydroxycoumarin **11**. Specifically, azide **11** was reacted with *endo*-BCN-OH **7** or *exo*-BCN-CH₂OH **8** under pseudo-first order conditions in MeOH/H₂O (1 : 1) to form the fluorescent triazole product (Scheme 2).²⁷ We found that *endo*-BCN-CH₂OH **7** reacted with a k_2 value of $0.90 (\pm 0.08) \text{ M}^{-1} \text{ s}^{-1}$ and *exo*-BCN-CH₂OH **8** with a k_2 value of $0.88 (\pm 0.06) \text{ M}^{-1} \text{ s}^{-1}$, resulting in a $k_{\text{rel},\text{endo/exo}}$ of 1.02 (see SI, Appendix E). Therefore, we were not able to confirm the reported difference between *exo* and *endo*-BCN-CH₂OH in SPAAC click chemistry.²⁸ Together with our analogous observation on the SPOCQ reaction, we conclude that it is likely that the reported difference in reactivity of *endo*-BCN and *exo*-BCN did not originate from intrinsic differences in reactivities of the two isomers of BCN. We conclude that the stereochemistry of the annulation is an ineffective approach to influence the reaction rate of SPAAC and SPOCQ reactions.

Crystal structure of *endo*-BCN-CH₂OH (**7**) and DBCO (**10**)

Structural details for the BCN and DBCO skeletons were obtained by X-ray diffraction on single crystals. Suitable crystals of *endo*-BCN-OH (**7**) were grown from a concentrated hot Et₂O solution, for DBCO (**10**) this was from a hot THF solution. The

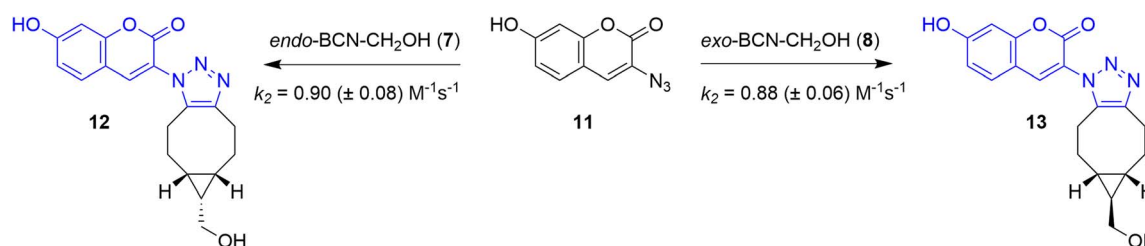
resulting crystal structures are depicted in Fig. 3a and b; unfortunately, we were not able to grow crystals from *exo*-BCN-CH₂OH (**8**).

Regarding the structure of *endo*-BCN-CH₂OH (**7**) (Fig. 3a), the endocyclic cyclopropane bond length is 1.52 Å, which is typical for sp³-sp³ hybridized carbon atoms. The structure reveals that ring pinching is afforded by means of reducing the interatomic distance between the two methylene groups that are attached to the cyclopropane ring and those next to the alkyne bond, *i.e.*, C14-C19 and C4-C9 to 3.277 Å and 3.294 Å, respectively (see SI, Appendix F). These interconnecting methylene groups of the propargylic position bend inwards to 105.2°–105.5° and of those next to the cyclopropane ring outwards to 112.0°–112.8°, which deviate substantially from 109.3° bond angles expected for sp³ carbon atoms,²⁹ forming a strained plane within the molecule. This forces BCN into a conformation in which the alkyne is oriented in a coplanar fashion (only 1.2° torsional angles, see Fig. 3a, structure on the right), resulting in bond angles of 154.2°–154.5° between the sp and sp³ hybridized carbon atoms. These angles are markedly narrower than the value of 158.5° that was reported for the cyclooctyne parent compound.³⁰ Therefore, the enhanced reactivity of BCN over that of cyclooctyne is caused by an increase in angular strain of the alkyne.^{19,31}

Inspection of the crystal structure of DBCO-acid **10** (Fig. 3b) reveals that the two individually planar aromatic rings are tethered together by a heavily distorted alkyne functionality, which is unsymmetrically bent at 152.3° and 154.2° (see SI, Appendix G). The alkyne possesses dihedral distortion by means of a 17.0° torsional angle, resulting in an antiparallel positioning of the aromatic planes relative to each other. The other bridge between the two rings is formed by a nitrogen atom and a methylene group, displaying a dihedral angle of 121.6°. Whereas the methylene group has a bond angle of 115.6°, thereby deviating from normal sp³ geometry, the nitrogen atom has an endocyclic bond angle of 120.2°, confirming sp² hybridization of this amide (sum of all bond angles around the endocyclic nitrogen atom is 359.6°). Interestingly, the attachment to the amide protrudes in an axial fashion from the DBCO ring system.

Alkyne bond angle and reactivity correlation

The XRD structures of *endo*-BCN-CH₂OH (**7**) and DBCO (**10**) (Fig. 3a and b) and of the earlier published structure of THS (**9**)¹⁷ allow assessment of the importance of bond angles of the



Scheme 2 Reaction of *endo*-BCN-CH₂OH **7** and *exo*-BCN-CH₂OH **8** with 3-azido-7-hydroxycoumarin **11** to form the fluorescent SPAAC products **12** and **13**, respectively, proceeds with identical second-order rate constants.



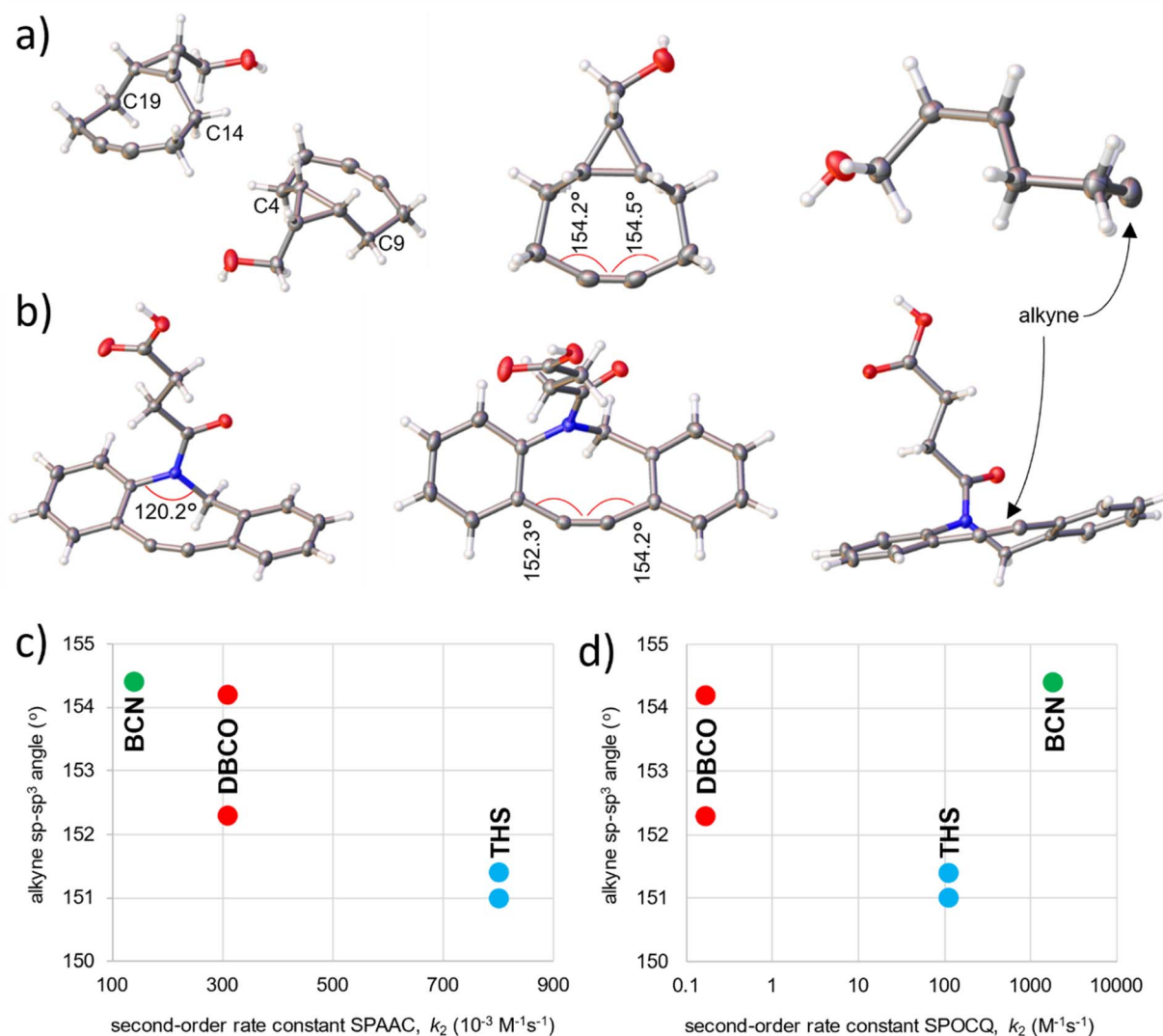


Fig. 3 (a) The molecular structure of *endo*-BCN-CH₂OH **7** according to X-ray structure determination. ORTEP depicted with thermal ellipsoids drawn at 50% probability level. The panel shows the presence of two molecules in the unit cell (left), and top and side view of isolated BCN molecules (middle and right); atoms and bond angles that are mentioned in the text are indicated. (b) The molecular structure of DBCO-acid **10** according to X-ray structure determination. ORTEP depicted with thermal ellipsoids drawn at 50% probability level. The panel shows DBCO from the side, top and front (left-to-right); bond angles that are mentioned in the text are indicated. (c) Chart showing a correlation between the strained alkyne bond angles, and the second-order rate constant for the SPAAC reaction ($R^2 = 0.9875$). (d) Chart showing the absence of a correlation between the strained alkyne bond angles, and the second-order rate constant for the SPOCQ reaction ($R^2 = 0.5470$). The ORTEP figures were generated with Olex2.³⁴

involved disubstituted alkynes in SPAAC and SPOCQ click reaction kinetics. THS possesses symmetrical alkyne bond angles of 151.0°–151.4° and is therefore the most strained stable and structurally well-characterized cycloalkyne to date.¹⁷ As described before, those of *endo*-BCN-CH₂OH (**7**) are 154.2°–154.5° and those of DBCO (**10**) are 152.3° and 154.2°. Plotting these angles to reported second-order rate constants for [3 + 2] SPAAC reveals a correlation with $R^2 = 0.9875$ (Fig. 3c).³² Such nice correlation is not found for the [4 + 2] SPOCQ reactions, where this plot shows $R^2 = 0.5470$ (Fig. 3d). It appears that the planar TS of [3 + 2] SPAAC – in which linear azide and alkyne form a planar triazole – fundamentally differs from the three-dimensionally demanding TS of [4 + 2] SPOCQ reactions as

the flat ring of the quinone is converted into an out-of-plane bent structure of the resulting bicyclo[2,2,2]octadiene. Our data shows that the height of the TS barrier in SPOCQ is primarily dictated by entropy and that reagent bulkiness has to be taken into account. For example, we previously ascribed the lower reactivity of THS with respect to BCN in SPOCQ to the presence of substantial steric bulk originating from the methyl groups neighboring the alkyne functionality.¹⁷ A similar reasoning can be applied for DBCO (**10**), which is even less reactive, as the presence of the aromatic rings that flank the alkyne in DBCO hinder its interaction with the quinone diene, suppressing the reaction rate constant.

Following Hammond's Postulate, structures of TSs often resemble the molecular structures of the reagents in the case of an early TS exergonic reaction. This means that the net bond breaking plus bond formation energy barrier (ΔH^\ddagger) would be lower for more strained systems, which is also seen in our dataset, *e.g.*, by comparing THS and *endo*-BCN-CH₂OH ($\Delta\Delta H^\ddagger = 1.4 \text{ kcal mol}^{-1}$).

A similar, but distinct phenomenon is observed for *trans*-cyclooctenes. The crown-like conformation of the molecule enforces dihedral bending of the $\text{sp}^2\text{-sp}^2$ bond plane of the alkene to an extent of 133.0° , as was shown crystallographically for an *O*-alkylated version of TCO-OH (**6**) by Fox *et al.* (Fig. 4a).³³ This 47° geometrical distortion of alkene planarity likely contributes strongly to the extremely low ΔH^\ddagger values found for cycloalkene SPOCQ reactions (Table 1). Annulation of the cyclopropane ring to TCO raises the energy of the structure by forcing it into a "half-chair"-conformation (Fig. 4a). In the absence of crystal structure information, the effect of this on the geometrical distortion of the alkene planarity is not known. However, as the activation enthalpy of both *s*TCO-CH₂OH structures, *i.e.*, **2** and **3**, is statistically indifferent from that of TCO, it can be assumed that a similar distortion of the double bond is also present in the *s*TCO family of dienophiles. Thus, the presence of the cyclopropane ring likely does not distort the alkene more than is already the case in TCO, but has a larger

effect of the conformation of the ring structure, as is also apparent from the large differences in the entropies of activation between TCO and the *s*TCO family, *i.e.*, $-15.5 \text{ kcal mol}^{-1}$ for TCO-OH (**6**) *vs.* $-11.9 \text{ kcal mol}^{-1}$ for both *endo*-*s*TCO-CH₂OH (**2**) and *exo*-*s*TCO-CH₂OH (**3**).

Lastly, we observed that the ΔH^\ddagger for *s*TCO-CH₂OH are slightly but significantly lower than those for the BCN-CH₂OH counterpart, with $0.6\text{--}1.0 \text{ kcal mol}^{-1}$ *versus* $1.3\text{--}2.5 \text{ kcal mol}^{-1}$ (standard deviations are included in these values), respectively. As such, involvement of secondary orbital interactions (SOIs) are not reflected in the enthalpy of activation in SPOCQ (Fig. 4b and c).

Computational investigations

The rather low activation enthalpies –which were also confirmed by computational analysis (see SI, Table S20)– also suggest that our dienophiles do not show a direct link between reaction rate and ring strain. To verify this, high-end calculations of the ring strain in the dienophiles (at the wb97M-V/def2-QZVPP//M06-2X/6-311+G(d,p) level of theory) were performed by calculating the free energy of ring-opened analogues of the ring structures under current study and those of the cycloalkenes and cycloalkynes for which we have activation parameters. These values were then compared to those of *n*-hexane and cyclohexane, as ring-shaped and ring-opened analogs that both did not display ring strain (see SI for details). These calculations show that ring strain is higher for the more reactive *trans*-cyclooctenes than for the less reactive cyclooctynes, and that this parameter varies gradually over the range of compounds within a range of *ca.* 5 kcal mol^{-1} (Table 2). However, there is no correlation between the ring strain and the activation enthalpy or activation free energy, and thus not with the reaction rate. In other words: in all these systems there is already sufficient ring strain to pre-load the reactants with such a drive that not enthalpy but entropy determines the reactions rates. Interestingly, the 7-membered THS **9** has, despite its smaller C–C≡C bond angles, an intermediate value for its ring strain, and not reaching those of our *s*TCO derivatives, which are clearly more strained than cyclooctynes like BCN. Comparison of the calculated ring strain with heats of hydrogenation calculated by the same method indeed reveals that THS **9** is in reactivity more

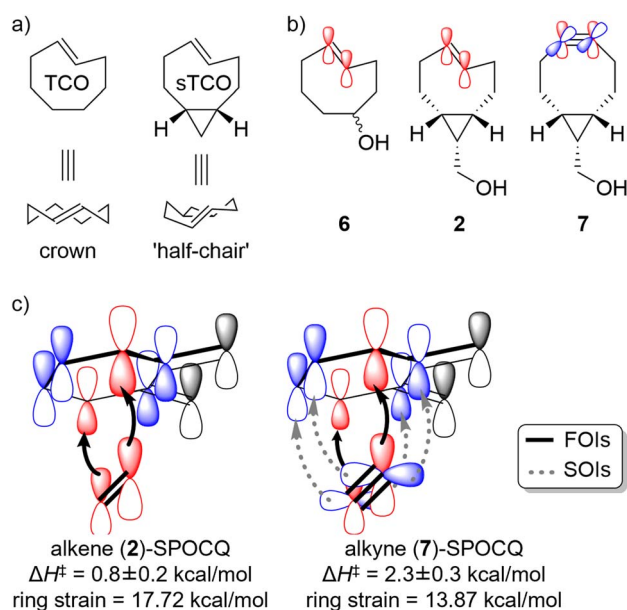


Fig. 4 Structural considerations regarding the interpretation of the thermodynamic activation parameters. (a) Effect of annulated cyclopropane ring on the conformation of *trans*-cyclooctene. (b) Graphical representation of the p-orbitals of the C=C bonds in TCO-OH (**6**) and *endo*-*s*TCO-CH₂OH (**2**) and of the C≡C bond in *endo*-BCN-CH₂OH (**7**). (c) Schematic depiction of the frontier orbital interactions (FOIs) of the alkene-SPOCQ and alkyne-SPOCQ reactions (between red p-orbitals, indicated by black arrows), and of proposed secondary orbital interactions (SOIs) of the alkyne-SPOCQ reaction (between the blue orbitals, indicated by grey arrows). Oxygen atoms in the diene and the ring of the dienophile are omitted for clarity.

Table 2 Ring strain calculations of different strained dienophiles **2–9** at the wb97M-V/def2-QZVPP//M06-2X/6-311+G(d,p) level of theory

Dienophile	Ring strain (kcal mol ⁻¹)
Cycloalkenes	
<i>endo</i> - <i>s</i> TCO-CH ₂ OH 2	17.72
<i>exo</i> - <i>s</i> TCO-CH ₂ OH 3	17.25
<i>endo</i> - <i>s</i> TCO-C(O)OEt 4	18.26
<i>exo</i> - <i>s</i> TCO-C(O)OEt 5	17.60
TCO-OH 6	18.41
Cycloalkynes	
<i>endo</i> -BCN-CH ₂ OH 7	13.87
<i>exo</i> -BCN-CH ₂ OH 8	14.53
THS 9	16.73

similar to the cycloalkenes **2–6** than to the cycloalkynes BCN-CH₂OH **7** and **8** (see SI, Fig. S3).

Conclusions

Temperature-dependent kinetic studies using stopped-flow UV-Vis spectroscopy provided detailed insights into *o*-quinone-cycloalkene click chemistry. We found that the ΔH^\ddagger values amongst the TCO dienophiles are all extremely small (<2 kcal mol^{−1}) with minimal differences within the series, and that mainly the activation entropy ΔS^\ddagger dictates the height of the ΔG^\ddagger barrier. Introduction of a cyclopropane moiety opposite to the dienophile in the eight-membered ring enforces reactivity and reduces the entropic barrier. We substantiated the reactivity of axial-TCO-OH and equatorial-TCO-OH with measured k_2 values of 34.8 M^{−1} s^{−1} and 2.8 M^{−1} s^{−1}, respectively, a 12.5-fold difference in reactivity in SPOCQ with quinone **1** in favour of the axial diastereomer. This settles the sTCO class of dienophiles as the new benchmark reagents in SPOCQ chemistry, reaching k_2 values of 3.5×10^3 M^{−1} s^{−1} for the alcohol-functionalized version **2** and **3**. Our reaction kinetics studies in MeOH reveal a strong on water-effect on the kinetics, as the second-order rate constants are approximately two orders of magnitude lower than in water-MeOH, 1 : 1 (v/v). Eyring analysis in MeOH revealed a similar value for the enthalpy of activation, but a substantially larger entropic barrier when compared to the water-MeOH mixture. Furthermore, no reactivity differences were found between *exo* and *endo* diastereomers of the probes in SPOCQ, unless bulkier substituents were positioned opposite of the dienophile. The long-claimed difference between *endo*-BCN-CH₂OH (**7**) and *exo*-BCN-CH₂OH (**8**) in SPAAC chemistry was also found not to be true. Furthermore, XRD analysis of both BCN and DBCO allowed us to conclude that the higher angular tension of DBCO does not affect its reactivity in SPOCQ, as BCN is 11 000 times more reactive, but that the bond angles of the disubstituted alkynes do correlate to the reaction rates observed for SPAAC. Finally, high-level DFT calculations yield as order of ring strains: sTCOs (17–18 kcal mol^{−1}) > cycloheptyne THS (16 kcal mol^{−1}) > BCN derivatives (14–15 kcal mol^{−1}). The absent correlation between ring strain and ΔH^\ddagger or ΔG^\ddagger confirms that the ring strain in these compounds is already sufficient to reduce the ΔH^\ddagger to insignificantly small values, and that ‘reducing chaos’ determines the reaction rate in SPOCQ chemistry. By uncovering the thermodynamic activation parameters of this SPOCQ click reaction, determining the crystal structure of the compounds, and calculations of the ring strain, we were able to construct a theoretical framework that enable future developments to push this unique set of biogenic click reactions into untapped domains, especially when high reaction rates are required.²

Methods

Stopped-flow UV-Vis kinetic studies

The reaction of 4-*tert*-butyl-*ortho*-quinone **1** with dienophiles **2–10** was measured under pseudo-first order conditions in 1 : 1 MeOH/MilliQ water following the decay of the specific

absorption band at 395 nm for *o*-quinone **1**. Two respective equivolume solutions of *o*-quinone **1** and the probe of interest, in MeOH : water – 1 : 1 (v/v) or in pure MeOH, were loaded into the two separate driver syringes of the RX2000 Rapid Kinetics Spectrometer Accessory (Applied Photophysics). The accessory is attached to a thermostat bath and to a Cary 60 UV-Vis spectrophotometer. The solutions in the driver syringes were thermostatted for at least 15 min prior to measurement. Upon measurement, the contents of the two driver syringes were flown simultaneously through the cuvette and measurement starts upon abruptly stopping the flow. Single wavelength measurements were then recorded every 12.5 ms at 395 nm. The measurements were performed in quadruplicate until the signal stabilizes. This setup utilizes equal volumes of the reagents, thereby halving each respective concentration in the cuvette. Concentrations are hereby referred to as final concentrations in the reaction mixture. The experiments were conducted using 40 μ M solutions of *o*-quinone **1** (1 eq) and 0.4–4 mM solutions of dienophile **2–9** (*i.e.*, 10–100 eq.) to allow for acquisition of sufficient data points for analysis. From these, k_2 plots were determined at 25 °C with the varying stoichiometry of the target probes. Eyring plots were determined at a set stoichiometry of 1 : 10 at varying temperatures of 5, 13, 21, 29, 37 °C. Measurements for DBCO-acid **10** (as potassium salt) were performed at a higher concentration of 4 mM DBCO with 0.4 mM *o*-quinone **1** at a set stoichiometry of 1 : 10 equivalents. Data analysis was then performed in GraphPad Prism 9 Version 9.3.1 (471) by exponential one phase decay fitting using nonlinear regression until a plateau of constant value is reached, leading to an observed pseudo-first-order rate constant k' (see SI for additional details). The k_2 values were then determined from the slope of the linear k' versus [dienophile] plot. The thermodynamic activation parameters ΔH^\ddagger and ΔS^\ddagger were determined *via* the classic method of Eyring utilizing the following linearized equation, with transmission coefficient κ (equals one); Boltzmann constant k_B ; Planck's constant h ; gas constant R ; temperature T (in K).^{35–39}

$$\ln \frac{k_2}{T} = \frac{-\Delta H^\ddagger}{R} \times \frac{1}{T} + \ln \frac{\kappa \times k_B}{h} + \frac{\Delta S^\ddagger}{R}$$

Fluorescence spectroscopy kinetic studies

Fluorescence measurements were performed on a Edinburgh Instruments FLS900 Fluorescence spectrometer at 20 °C under pseudo-first order conditions in 1 : 1 MeOH/MilliQ water (v/v). Two equivolume solutions of 10 μ M 3-azido-7-hydroxycoumarin **11** and 10 mM *endo*-BCN-CH₂OH **7** or *exo*-BCN-CH₂OH **8** were mixed in a quartz cuvette at a set stoichiometry of 1 : 1000 equivalents. Concentrations are referred to as final concentrations as in the reaction mixture. Formation of the fluorescent triazole click product was followed over time (λ_{ex} = 395 nm; λ_{em} = 472 nm). Emission spectra were recorded every 9 seconds for 15 minutes at 472 nm, at which point the increase of signal reached a plateau. Data analysis was then performed in GraphPad Prism 9 Version 9.3.1 (471) by exponential plateau fitting using nonlinear regression, leading to an observed pseudo-first-order rate constant k' , from which the k_2 values



were then obtained by dividing k' by [*endo*-BCN-CH₂OH] or [*exo*-BCN-CH₂OH]. The measurements were performed in triplicate for each compound.

Abbreviations

ax	Axial
BCN	Bicyclo[6.1.0]non-4-yne
DBCO	Aza-dibenzocyclooctyne
eq	Equatorial
ΔG^\ddagger	Gibbs energy of activation
ΔH^\ddagger	Enthalpy of activation
ΔS^\ddagger	Entropy of activation
k_2	Second-order rate constant
SOI	Secondary orbital interaction
SPAAC	Strain-promoted (3 + 2) azide-alkyne cycloaddition
SPOCQ	Strain-promoted oxidation-controlled <i>ortho</i> -quinone cycloaddition
sTCO	Strained- <i>trans</i> -cyclooctene
TCO	<i>trans</i> -Cyclooctenol
THS	3,3,6,6-Tetramethyl-1-thiacyclo-heptyne sulfoximide
TS	Transition state

Author contributions

J. A. M. D. performed most experiments in water-MeOH and grew the single crystals for X-ray analysis, J. F. performed all experiments in MeOH, J. E. and H. Z. performed the computational analysis, B. A. conceived the project, J. A. M. D. and B. A. wrote the manuscript, all authors delivered input for the manuscript and approved the final version.

Conflicts of interest

The authors declare no competing financial interest.

Data availability

Crystallographic data for structures of compounds **7** and **10** have been deposited at the Cambridge Crystallographic Data Centre (CCDC) under deposition numbers 2378937 for *endo*-BCN-CH₂OH (**7**) and 2378938 for DBCO (**10**).

CCDC 2378937 and 2378938 contain the supplementary crystallographic data for this paper.^{40a,b}

The data supporting this article have been included as part of the supplementary information (SI). Supplementary information: experimental procedures; kinetic data; analyses; X-ray crystallographic data of *endo*-BCN-CH₂OH **7** and DBCO-acid **10**; and computational studies and cartesian coordinates. See DOI: <https://doi.org/10.1039/d5sc04275e>.

Acknowledgements

We thank Dr. Gabriele Kociok-Köhn (University of Bath, UK) for her work on obtaining and processing of the single crystal X-ray structures, and Barend van Lagen for his assistance with

fluorescence spectroscopy. This work is funded by the NWO Gravity Program Institute for Chemical Immunology (ICI), project number ICI00027 (J. A. M. D. and B. A.). H. Z. thanks Jiaying University for generous support.

References

- 1 A. Borrmann, O. Fatunsin, J. Dommerholt, A. M. Jonker, D. W. P. M. Löwik, J. C. M. Van Hest and F. L. van Delft, *Bioconjugate Chem.*, 2015, **26**, 257–261.
- 2 B. Albada, J. F. Keijzer, H. Zuilhof and F. van Delft, *Chem. Rev.*, 2021, **121**, 7032–7058.
- 3 J. J. Bruins, A. H. Westphal, B. Albada, K. Wagner, L. Bartels, H. Spits, W. J. H. van Berkel and F. L. van Delft, *Bioconjugate Chem.*, 2017, **28**, 1189–1193.
- 4 J. J. Bruins, B. Albada and F. van Delft, *Chem.-Eur. J.*, 2018, **24**, 4749–4756.
- 5 J. J. Bruins, D. Blanco-Ania, V. van der Doef, F. L. van Delft and B. Albada, *Chem. Commun.*, 2018, **54**, 7338–7341.
- 6 J. J. Bruins, C. van de Wouw, J. F. Keijzer, B. Albada and F. L. van Delft, in *Enzyme-Mediated Ligation Methods*, ed. T. Nuijens and M. Schmidt, Springer New York, New York, NY, 2019, pp. 357–368.
- 7 J. J. Bruins, C. van de Wouw, K. Wagner, L. Bartels, B. Albada and F. L. van Delft, *ACS Omega*, 2019, **4**, 11801–11807.
- 8 J. J. Bruins, J. A. M. Damen, M. A. Wijdeven, L. P. W. M. Lelieveldt, F. L. van Delft and B. Albada, *Bioconjugate Chem.*, 2021, **32**, 2167–2172.
- 9 I. Shajan, L. N. C. Rochet, S. R. Tracey, B. Jackowska, R. Benazza, O. Hernandez-Alba, S. Cianférani, C. J. Scott, F. L. van, D. V. Chudasama and B. Albada, *Bioconjugate Chem.*, 2023, **34**, 2215–2220.
- 10 A. M. Jonker, A. Borrmann, E. R. H. van Eck, F. L. van Delft, D. W. P. M. Löwik and J. C. M. van Hest, *Adv. Mater.*, 2015, **27**, 1235–1240.
- 11 R. Sen, J. Escorihuela, F. van Delft and H. Zuilhof, *Angew. Chem., Int. Ed.*, 2017, **56**, 3299–3303.
- 12 D. Gahtory, R. Sen, A. R. Kuzmyn, J. Escorihuela and H. Zuilhof, *Angew. Chem., Int. Ed.*, 2018, **57**, 10118–10122.
- 13 J. Escorihuela, A. Das, W. J. E. Looijen, F. L. van Delft, A. J. A. Aquino, H. Lischka and H. Zuilhof, *J. Org. Chem.*, 2018, **83**, 244–252.
- 14 J. Escorihuela, W. J. E. Looijen, X. Wang, A. J. A. Aquino, H. Lischka and H. Zuilhof, *J. Org. Chem.*, 2020, **85**, 13557–13566.
- 15 B. J. Levandowski, D. Svatunek, B. Sohr, H. Mikula and K. N. Houk, *J. Am. Chem. Soc.*, 2019, **141**, 2224–2227.
- 16 B. J. Levandowski, D. Svatunek, B. Sohr, H. Mikula and K. N. Houk, *J. Am. Chem. Soc.*, 2019, **141**, 18641.
- 17 J. A. M. Damen, J. Escorihuela, H. Zuilhof, F. L. van Delft and B. Albada, *Chem.-Eur. J.*, 2023, **29**, e202300231.
- 18 M. T. Taylor, M. L. Blackman, O. Dmitrenko and J. M. Fox, *J. Am. Chem. Soc.*, 2011, **133**, 9646–9649.
- 19 J. Dommerholt, S. Schmidt, R. Temming, L. J. A. Hendriks, F. P. J. T. Rutjes, J. C. M. van Hest, D. J. Lefeber, P. Friedl and F. L. van Delft, *Angew. Chem., Int. Ed.*, 2010, **49**, 9422–9425.



- 20 A. Darko, S. Wallace, O. Dmitrenko, M. M. Machovina, R. A. Mehl, J. W. Chin and J. M. Fox, *Chem. Sci.*, 2014, **5**, 3770–3776.
- 21 M. Wang, D. Svatunek, K. Rohlfing, Y. Liu, H. Wang, B. Giglio, H. Yuan, Z. Wu, Z. Li and J. Fox, *Theranostics*, 2016, **6**, 887–895.
- 22 A. George, G. Krishna Priya, M. Ilamaran, N. R. Kamini, S. Ganesh, S. Easwaramoorthi and N. Ayyadurai, *ChemistrySelect*, 2017, **2**, 7117–7122.
- 23 A. George, M. Indhu, S. Ashokraj, P. Shanmugam, N. R. Ganesan, N. R. Kamini and N. Ayyadurai, *Bioorg. Med. Chem.*, 2021, **50**, 116460.
- 24 H. Zhan, H. de Jong and D. W. P. M. Löwik, *ACS Appl. Bio Mater.*, 2019, **2**, 2862–2871.
- 25 The difference in reactivities and difference in the height of the activation barriers is in accordance with $\Delta\Delta G^\ddagger = -R \cdot T \cdot \ln(k_{\text{rel}})$.
- 26 A. Lubineau, H. Bienaymé, Y. Queneau and M. Scherrmann, *New J. Chem.*, 1994, **18**, 279.
- 27 K. Sivakumar, F. Xie, B. M. Cash, S. Long, H. N. Barnhill and Q. Wang, *Org. Lett.*, 2004, **6**, 4603–4606.
- 28 J. Dommerholt, O. van Rooijen, A. Borrmann, C. F. Guerra, F. M. Bickelhaupt and F. L. van Delft, *Nat. Commun.*, 2014, **5**, 5378.
- 29 M. B. Smith and J. March, *March's Advanced Organic Chemistry: Reactions, Mechanisms, and Structure*, Hoboken, New Jersey, United States of America, John Wiley & Sons, Inc., 6th edn, 2007.
- 30 J. Haase and A. Krebs, *Z. Naturforsch. A*, 1971, **26**, 1190–1193.
- 31 C. Antony-Mayer and H. Meier, *Chem. Ber.*, 1988, **121**, 2013–2018.
- 32 C. G. Gordon, J. L. MacKey, J. C. Jewett, E. M. Sletten, K. N. Houk and C. R. Bertozzi, *J. Am. Chem. Soc.*, 2012, **134**, 9199–9208.
- 33 M. Royzen, G. P. A. Yap and J. M. Fox, *J. Am. Chem. Soc.*, 2008, **130**, 3760–3761.
- 34 O. V. Dolomanov, L. J. Bourhis, R. J. Gildea, J. A. K. Howard and H. Puschmann, *J. Appl. Cryst.*, 2009, **42**, 339–341.
- 35 H. Eyring, *J. Chem. Phys.*, 1935, **3**, 107–115.
- 36 A. Dirksen and P. E. Dawson, *Bioconjugate Chem.*, 2008, **19**, 2543–2548.
- 37 M. G. Evans and M. Polanyi, *Trans. Faraday Soc.*, 1935, **31**, 875–894.
- 38 W. F. K. Wynne-Jones and H. Eyring, *J. Chem. Phys.*, 1935, **3**, 492–502.
- 39 K. J. Laidler and M. C. King, *J. Phys. Chem.*, 1983, **87**, 2657–2664.
- 40 (a) J. A. M. Damen, J. Escorihuela, J. Firet, H. Zuilhof and B. Albada, CCDC 2378937: Experimental Crystal Structure Determination, 2025, DOI: [10.5517/ccdc.csd.cc2kvgxt](https://doi.org/10.5517/ccdc.csd.cc2kvgxt); (b) J. A. M. Damen, J. Escorihuela, J. Firet, H. Zuilhof and B. Albada, CCDC 2378938: Experimental Crystal Structure Determination, 2025, DOI: [10.5517/ccdc.csd.cc2kvgyv](https://doi.org/10.5517/ccdc.csd.cc2kvgyv).

