

Cite this: *Chem. Sci.*, 2025, 16, 12242

Topological variety and self-sorting in homo- and heteroleptic Pd_nL_{2n} metallo-supramolecular assemblies

Laura Neukirch and Guido H. Clever *

A plethora of nanoscale Pd_nL_{2n}-type architectures has been synthesized through the coordination-driven self-assembly of Pd(II) ions and organic bis-monodentate bridging ligands. While initially, the focus was on homoleptic structures, comprising one type of ligand per assembly, the field has recently shifted towards reducing symmetry in heteroleptic multicomponent assemblies, containing two or more distinct ligands in defined positions. In parallel, the incorporation of functional moieties such as binding and catalytic sites, photoswitches and redox units has seen a steep development. While empirical data have been gathered on the relationship between the ligand structure and assembly outcome for a limited number of cases, confidently forecasting the result of reacting a given ligand with Pd(II) cations often still remains challenging and has been mastered only for the simplest systems. Additionally, new Pd_nL_{2n} topologies – along with subtle factors driving their formation (such as counter anion or guest templation and solvation effects) – are discovered continuously. For designing metallosupramolecular assemblies for application, it is of pivotal importance to increase predictability and gain control over assembly topology, as the structure and properties are often closely connected. To raise awareness for the problem's complexity, we commence this review by exploring the surprising breadth of topological diversity among homoleptic Pd_nL_{2n} ($n = 2-8$) architectures that has so far been found experimentally. We next discuss strategies for increasing the structural complexity even further through the non-statistical self-assembly of heteroleptic cages, the orientational self-sorting of asymmetric ligands, and

Received 2nd May 2025
Accepted 12th June 2025DOI: 10.1039/d5sc03203b
rsc.li/chemical-science

Department of Chemistry and Chemical Biology, TU Dortmund University,
Otto-Hahn-Str. 6, 44227 Dortmund, Germany. E-mail: guido.clever@tu-dortmund.de



Laura Neukirch

Laura Neukirch obtained her bachelor's degree in chemical biology from TU Dortmund University in 2019. Following predoctoral studies in chemistry, she joined the Clever Lab as a fast-track PhD student within the Cluster of Excellence RESOLV in 2021. Her research mainly focuses on Pd(II)-based donor-acceptor cages for photo-induced charge separation. This brought her to the Massachusetts Institute of Technology,

where she worked with the group of Prof. Schlau-Cohen for a few months. She also investigated the self-assembly of five-membered rings Pd₅L₁₀, which sparked her interest in strategies for accessing different Pd_nL_{2n} topologies that are described in this review article.



Guido H. Clever

Guido Clever has been a full professor at TU Dortmund University since 2015 and holds the chair of Bioinorganic Chemistry. He studied chemistry in Heidelberg and obtained his PhD in 2006 from LMU Munich. As a JSPS/Alexander-von-Humboldt fellow he then joined The University of Tokyo, becoming an assistant professor in 2009. In 2010 he became a junior professor at Göttingen University and an associate

professor in 2013. He received the ADUC award of the German Chemical Society, the FCI Dozenten Award and an ERC Consolidator grant. He is a board member of the Cluster of Excellence RESOLV, GRK2376 and journal Aggregate. His work focuses on non-statistical assembly, multi-chromophore architectures and metal-binding G-quadruplex DNA.



chiral self-sorting effects. Special emphasis will be placed on factors governing the particular self-assembly outcome as well as on rationalization approaches based on computations or geometrical considerations.

1. Introduction

Since the beginning of the field of supramolecular chemistry, researchers have aspired to understand and mimic the high selectivity and specificity of structure formation, binding processes and catalysed reactions found in natural systems.¹ A closer look into the nanosized cavities of enzymes reveals their often highly unsymmetric, multi-functionalized and dynamic nature.² While imitating certain aspects of the structural and functional complexity of natural systems by means of classical covalent synthesis is possible in principle,^{3,4} such approaches usually involve resource-consuming multi-step organic conversions and tedious purification tasks. The concept of dynamic self-assembly represents an attractive alternative to this: the combination of suitable molecular building blocks with a pre-programmed connectivity allows for the formation of large and intricate supramolecular architectures, usually under thermodynamic or sometimes kinetic control.⁵ Supramolecular assembly approaches can be classified according to the interactions that drive recognition and binding between the different building blocks. Popular approaches are based on dynamic covalent reactions,^{6–9} hydrogen bonding,^{10–13} and metal coordination.^{14–22} In the latter case, metal ions are combined with organic building blocks equipped with a set of suitable donor groups, serving as ligands to bridge two or more metal nodes. Often, transition metal cations with defined coordination geometries are used to enhance predictability of the self-assembly outcome. This approach allows for the construction of extended structures such as coordination polymers or metal organic frameworks (MOFs) as well as discrete, soluble entities such as defined rings, helicates and cages. This review only covers the latter class of non-polymeric compounds. Pioneering examples of such supramolecular coordination architectures include Saalfrank's Mg(II)-based adamantoid assembly from 1988,²³ Lehn's helicates,²⁴ Raymond's Ga(III) catecholate tetrahedra,²⁵ and Stang's Pt(II)-based rings and cages.²⁶ As a very versatile metal node, Pd(II) was found to assemble with nitrogen-containing heterocyclic donors such as pyridines and isoquinolines to form defined supramolecular architectures. Aside from the fixed square-planar coordination geometry of d⁸-configured Pd(II) cations, their diamagnetism (beneficial for NMR analytics), suitable ligand exchange kinetics and the thermodynamic stability of their *N*-ligand complexes render them highly attractive for realizing a large variety of metal-mediated assemblies.¹⁶

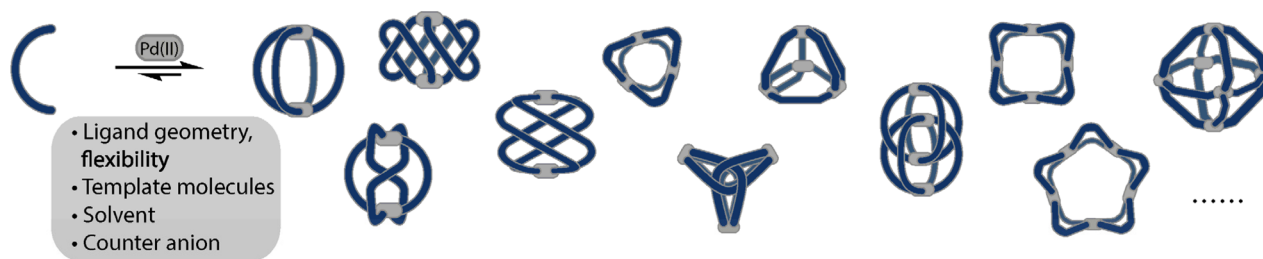
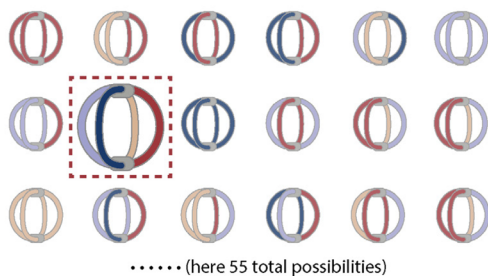
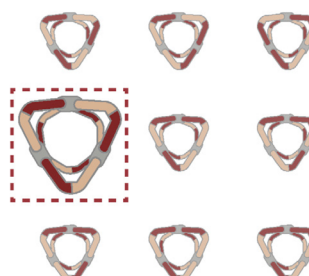
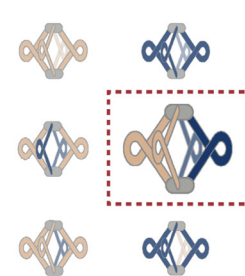
In the early 1990s Makoto Fujita pioneered the construction of two- and three-dimensional assemblies based on *cis*-protected Pd(II) cations and organic bridging panel-shaped ligands.^{20,27} In 1998, McMorran and Steel reported the first self-assembly of a helical Pd₂L₄ cage by employing a Pd(II) salt with a non-coordinating hexafluorophosphate counter anion and bridging bis-pyridyl ligands with a flexible backbone

structure.²⁸ Since these initial studies, interest in constructing and applying such Pd(II) assemblies blossomed.^{21,29–31} Incipiently, the focus of the field was mainly on reporting the successful design or serendipitous finding of new coordination assemblies, largely driven by their aesthetic appeal and the quest for unprecedented topologies. In the last decade, the field transitioned towards implementing functionality into coordination cages, thus bringing them closer to application. In this direction, strategies for modular multicomponent self-assembly started to be developed, allowing the arrangement of several different organic building blocks in close proximity without the need for extensive organic synthesis. Examples for functions that have been incorporated into Pd_{*n*}L_{2*n*} assemblies include catalytic activity,^{32–37} light harvesting chromophores,³⁸ fluorophores,^{39–42} (photo)-redox activity,^{43–45} photoswitches,^{46–50} and solubility-controlling groups for higher-order self-assembly.^{51–53} Furthermore, many assemblies possess a cavity in which guest molecules can be encapsulated, allowing them to serve as molecular reaction vessels^{32,54–56} and high-affinity receptors,^{57–59} among other functions.⁶⁰ However, most architectures reported so far have high symmetry, sharply contrasting with the low symmetry environments found in natural host systems.^{61–63} Therefore, efforts have been devoted to lowering the symmetry of such artificial self-assemblies by the controlled use of (I) matching organic ligands to form heteroleptic cages,^{21,64} (II) asymmetric ligands,³⁰ and (III) different metal-ions to form heteronuclear assemblies.⁶⁵ Self-sorting has also been investigated in the context of organic cages.⁶⁶

Despite vast strides being made in this field, many of the rules governing the formation of specific and sometimes surprising topologies – often closely connected to the functionality of the assemblies – still remain elusive. Thus, we begin this review by exploring the various topologies found for Pd_{*n*}L_{2*n*} homoleptic assemblies, *i.e.* architectures constructed from only one kind of organic ligand (Scheme 1a). While self-assembly is an equilibrium reaction, for simplicity, we use simple arrows throughout the remaining review. Next, we set out to summarise strategies developed for achieving integrative self-sorting to defined heteroleptic assemblies (Scheme 1b). We will then discuss approaches for orientational self-sorting to form homoleptic cages based on asymmetric ligands (Scheme 1c). Lastly, we describe the formation of cages *via* chiral self-sorting phenomena (Scheme 1d).

While there are many reports on assemblies with *cis*- or *trans*-protected Pd(II) cations and assemblies with ligands of higher denticity^{15,21} as well as on heteronuclear cages,⁶⁵ we limit our discussion here to architectures based on bis-monodentate organic ligands and unprotected Pd(II) cations. We start this review by highlighting the breadth of topological diversity among homoleptic Pd_{*n*}L_{2*n*} (*n* = 2–8) architectures. We then discuss strategies for the non-statistical self-assembly of heteroleptic cages, followed by orientational self-sorting of



a) **Topology control** in homoleptic assembliesb) **Integrative** self-sorting in heteroleptic cagesc) **Orientalional** self-sortingd) **Chiral** self-sorting

Scheme 1 Emergence of structural diversity in Pd_nL_{2n} coordination assemblies. (a) Examples for homoleptic architectures with different topologies and nuclearities (the inset shows some factors that may influence the outcome). (b) Integrative self-sorting to one defined heteroleptic cage out of numerous possible dinuclear structures based on four different ligands. (c) Orientalional self-sorting to one isomer of a three-membered ring from an asymmetric ligand. (d) Self-sorting of a racemic mixture of chiral ligands to a single cage isomer.

asymmetric ligands and chiral self-sorting. Special emphasis will be placed on factors governing the particular self-assembly outcome and attempts to rationalize experimental results based on computations or geometrical considerations.

2. Homoleptic dinuclear assemblies ($n = 2$)

The most basic and best explored family member among Pd_nL_{2n} cages consists of two $\text{Pd}(\text{II})$ ions bridged by four organic ligands. After the initial work on a quadruple stranded helicate Pd_2L_4 by McMorran and Steel,²⁸ a multitude of cages and helicates has been reported, with many studies focusing on the creation of larger cavities^{67,68} for the encapsulation of big^{58,69,70} or multiple^{58,71,72} guests and/or towards the incorporation of functionalities in the ligand backbone.^{38–40,42–47,50,51} As the field of the so-called lantern-shaped Pd_2L_4 cages has been reviewed recently,²⁹ we will only briefly discuss some key parameters governing their formation and highlight prominent examples. Some helicates will be tackled with special emphasis on the degree of helicity. Finally, we describe in more depth some recent examples for some rather sophisticated dinuclear topologies, including those that do not obey the principle of maximum site occupancy, *i.e.* assemblies of stoichiometry Pd_2L_2 and Pd_2L_3 .

2.1. Lantern-shaped cages and helicates

The geometry of rigid bispyridyl-type bridging ligands can be described through the binding angle α , defined as the angle

between the two nitrogen donor vectors. For ligands with collinear vectors, so-called lantern-shaped cages are preferably formed (Fig. 1a). A prominent example for a lantern-shaped cage is $\text{Pd}_2(\mathbf{1.1a})_4$, (note the naming scheme used herein: ligand **1.1a** reacts to form cage **1.2a**), initially reported by Hooley and coworkers in 2010.⁷³ Here, the π -surfaces of the ligands are oriented in a perpendicular fashion with respect to the longer

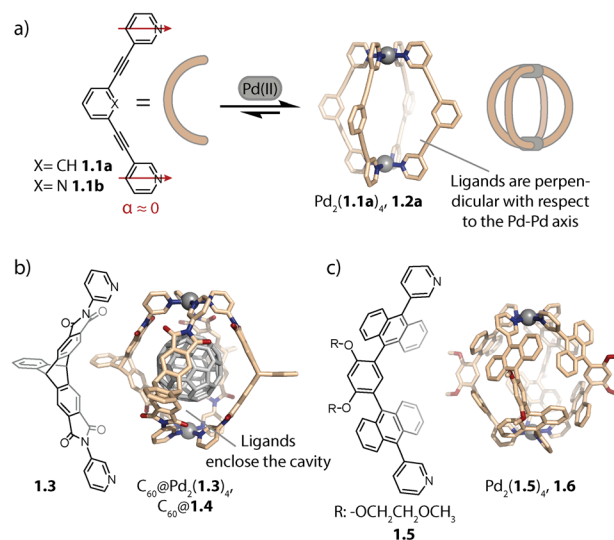


Fig. 1 (a) Self-assembly of lantern-shaped cage **1.2a**. (b) Capsule **1.4** based on a triptycene-derived ligand, encapsulating a C_{60} fullerene. (c) Capsule **1.6** based on ligands with anthracene panels.



axis of the cavity (the latter going through both Pd centres), allowing for inward-pointing functionalities to be placed in the backbones. Lusby and coworkers exploited the inward pointing α -protons of the coordinating pyridines for forming multiple hydrogen bonds to the oxygens of quinones, resulting in encapsulation with high association constants.⁷⁴ Furthermore, the same group used **1.2b** for efficiently catalysing cycloadditions to the thus activated substrates inside the cage cavity.^{33–35}

On the other hand, ligands with aromatic surfaces that partially enclose the cavity have been employed to assemble Pd₂L₄ capsules. For example, we designed curved ligand **1.3** based on triptycene^{58,75} (and ligands with a similarly shaped backbone⁷⁶) aiming at a size match between the cavity of the corresponding capsule Pd₂(**1.3**)₄, **1.4** and fullerene C₆₀ (Fig. 1b). This allowed selective binding of C₆₀ and thereby its solubilization in polar organic solvents.

Yoshizawa and co-workers reported capsule Pd₂(**1.5**)₄, **1.6** based on ligand **1.5**, whose methoxy ethoxy substituents endow the assembly with good solubility, even in water (Fig. 1c). The cavity of **1.6** is surrounded by an array of anthracene panels, creating a hydrophobic confinement. Owing to these two key properties, capsule **1.6** is capable of binding a range of neutral guests in aqueous solution.⁵⁹ Recently, the same group used a similar capsule decorated with sugar substituents to induce chirality on an encapsulated fullerene.⁷⁷ Our group, in collaboration with Gunnlaugsson and coworkers, has also achieved chirality transfer onto fullerenes (C₆₀ and C₇₀) with a readily accessible, relatively lightweight Pd₂L₄ cage with ligands constructed from Tröger's base.⁷⁸

Certain factors such as the ligand geometry,^{38,79,80} steric congestion at the backbone⁸¹ or donor site,⁸² and attractive inter-ligand interactions^{83,84} can favour a helical twisting of the ligands around the Pd–Pd axis. For characterizing the degree of helicity, the azimuthal angle β can be used. It is defined as the dihedral angle between the two N–Pd coordination bonds of the same ligand (Fig. 2a).

In their initial report, McMorran and Steel found one PF₆[–] anion encapsulated in the cavity of helicate Pd₂(**1.7**)₄ which is based on flexible ligand **1.7** (Fig. 2a).²⁸ This recently enticed the authors to study helicate formation in the presence of a variety of counter anions. Tetrahedral anions were found to be preferred over octahedral PF₆[–] and amongst the anions tested, ClO₄[–] showed the strongest binding, probably due to a perfect size match. On decreasing the size of the anions, the assembly shrinks along its Pd–Pd axis, concomitant with an increase in helicity, enabling an almost constant degree of cavity filling. The largest azimuthal angle of 85° was observed with iodine.⁸⁵

Chand and coworkers reported in 2010 on helicate Pd₂(**1.8**)₄ consisting of piperazine-based ligand **1.8**. The ligand adopts a flattened *anti* conformation upon self-assembly and the helicate is characterized by an azimuthal angle of around 45°.⁸⁶ In 2012, the same group reported a helicate with a similar azimuthal angle by employing another flexible ligand **1.9** based on a central pyridine (not participating in metal coordination) that was equipped with amide linkers.⁸⁷ The ligand can adopt different conformations depending on the orientation of the amide linkages relative to the backbone. It was found that the carbonyl bonds rotate outside of the helicate, presumably to

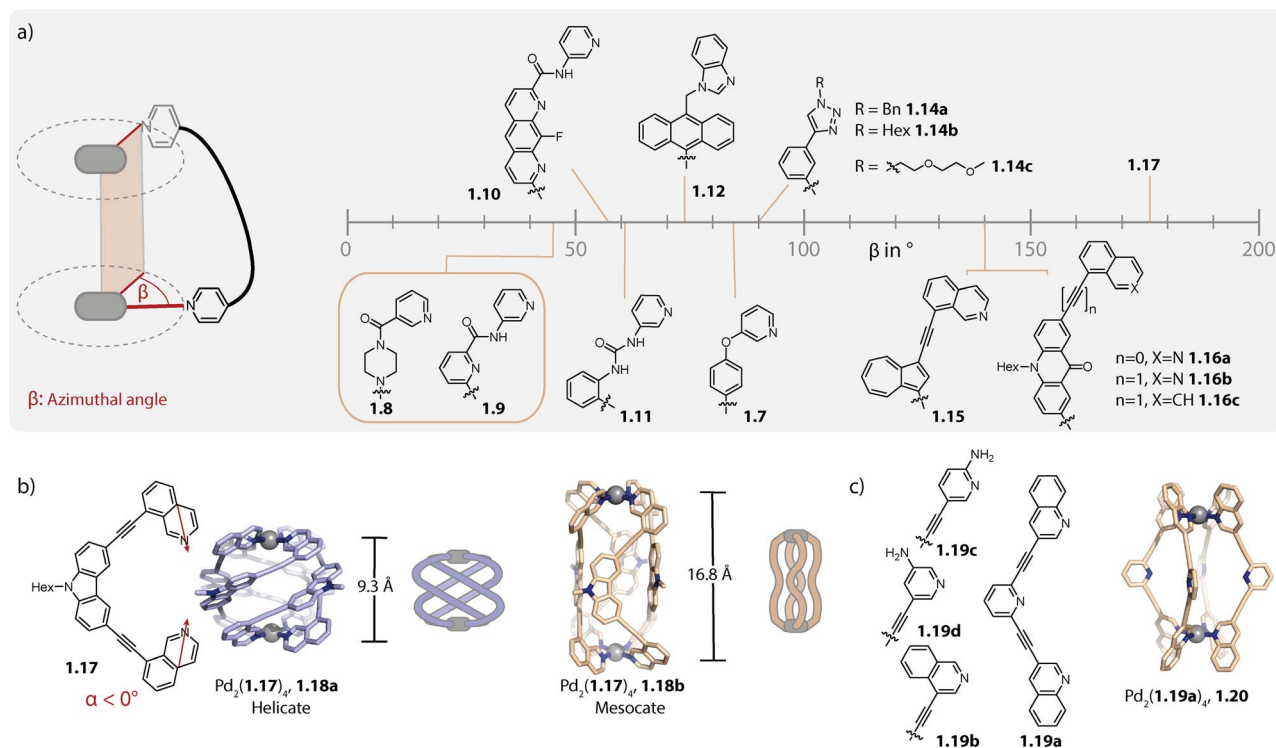


Fig. 2 (a) Overview of azimuthal angles β of some Pd₂L₄ helicates formed with different bis-monodentate organic ligands (the β values were taken from the corresponding publication when available or were determined based on the X-ray crystal structure as described in ref. 96). (b) Structure of helicate **1.18a** in comparison to isomeric mesocate **1.18b**. (c) Structure of ligands **1.19** and mesocate **1.20**.



avoid electronic repulsion with the central pyridine nitrogen and for enabling H-bonding between the NH groups and the encapsulated anion. In contrast, the same ligand with a benzene core formed, albeit under different conditions, a lantern-shaped cage, as reported by Puddephatt and coworkers in 2004. In this case, one of the two carbonyl groups of each ligand indeed points inside the cavity.⁸⁸

In 2018, Gan and coworkers showed that ligand **1.10** assembles to form a mixture of D_4 -symmetric helicate and C_{4h} -symmetric mesocate when combined with $\text{Pd}(\text{NO}_3)_2$. Selective formation of helicate or mesocate was favoured with a $\text{B}_{12}\text{F}_{12}^{2-}$ guest or with tetrafluoroborate, respectively. Furthermore, addition of the chiral anion Δ -TRISPHAT led to an enrichment of one of the helical enantiomers, as the authors showed by CD spectroscopy.⁸⁹

Additionally, in 2021, Chand and coworkers reported the anion-controlled self-assembly of helicate *versus* lantern-shaped cages: ligand **1.11** with urea linkers yields non-helical $\text{Pd}_2(\mathbf{1.11})_4$ in the presence of NO_3^- while a helicate with an angle of approximately 60° is formed with ClO_4^- anions.⁹⁰

Sun and coworkers reported in 2015 on helicate $\text{Pd}_2(\mathbf{1.12})_4$ comprising anthracene-based ligands **1.12** equipped with benzimidazole donor moieties. Ligand twisting, resulting in a D_4 -symmetrical helicate with an azimuthal angle of around 75° , can be traced back to the steric repulsion between the anthracene backbones. The latter creates a hydrophobic cavity which encapsulates a nitrate anion that is stabilized by multiple H-bonds to the acidic imidazole protons. The authors showed that the binding affinity towards nitrate is significantly higher as compared to those of other anions tested.⁸¹

Crowley and coworkers reported helicate $\text{Pd}_2(\mathbf{1.14a})_4$ based on a small triazine-terminated ligand modified with benzyl groups. The helicate is further stabilized by π - π interactions and the azimuthal angle reaches around 90° .⁹¹ In a follow-up study, the group modified the triazine appended moiety as well as the spacer. For the latter, the helicates were shown to form reliably when the 1,3-phenylene moiety was substituted with a flexible propyl chain but not when substituted with a 1,4-phenylene group. Furthermore, the helical structure displayed tolerance across all tested exohedral triazine substituents, ranging from alkyl chains to electron-rich and -poor aromatic groups.⁹² In the context of biological activity, the Crowley and Giles group studied the stability of $\text{Pd}_2(\mathbf{1.14a-c})_4$ (Fig. 2a) and **1.2b** (see Fig. 1a) against biological nucleophiles. Here, the helicates showed higher stability compared to lantern-shaped cage **1.2b**, presumably due to the increased donor strength of the triazine and the sterically shielded Pd(II) centres. Furthermore, $\text{Pd}_2(\mathbf{1.14b})_4$ proved more stable compared to $\text{Pd}_2(\mathbf{1.14a})_4$ and $\text{Pd}_2(\mathbf{1.14c})_4$, originating most likely from the more restricted access to the coordination sites, owing to agglomeration of the alkyl chains.⁹³

We have reported the self-assembly of helicates from various ligands possessing 8-isoquinoline donor moieties. The choice of this donor moiety leads to strongly convergent bonding vectors and hence, the ligands have to severely twist for bridging the two Pd(II) ions. The helical arrangement is further stabilized by π - π interactions between the ligands. Self-assembly of

ligands **1.15** (ref. 79) and **1.16a**, **1.16b** (ref. 80 and 94) which are based on azulene and acridone backbones, respectively, led to helicates with azimuthal angles of around 140° . We also synthesized helicates based on well-known coal-tar dyes; in this case, piperazine linkers were used, allowing for conserving the absorption properties of the dyes.³⁸ The corresponding ligands based on acridone, azulene, or coal-tar dyes equipped with 3-pyridyl donors instead of 8-isoquinolines assemble to form interlocked double cages⁹⁵ or to lantern-shaped cages,^{38,79} respectively. This can be traced back to the nearly collinear binding vectors of the latter, showcasing how the topology of the dinuclear species in terms of its helical twist can be controlled with the ligand geometry.

In order to attain even greater azimuthal angles, we recently investigated the self-assembly of carbazole-based ligand **1.17** with an even more negative binding angle. Formation of helicate $\text{Pd}_2(\mathbf{1.17})_4$, **1.18a** was observed with various different Pd(II) salts and the azimuthal angles ranged between 171 and 176° (Fig. 2b). In each case, one anion was found to be encapsulated within the cavity. Interestingly, with an increase in the size of anion, the Pd-Pd distance decreased, concomitant with a widening of the cage *i.e.* a change in the shape from prolate to oblate. While this might sound contradictory to the earlier report of McMorran and Steel,⁸⁵ the structural adaptation serves the same purpose, that is to retain a favourable cavity filling. Helicate **1.18a** exists in a counter anion- and solvent dependent equilibrium with the corresponding mesocate **1.18b**. In the latter, the ligands adopt C_s symmetry, resulting in a significantly larger Pd-Pd distance (16.8 \AA as compared to 9.3 \AA) and an overall achiral structure.⁹⁶

Crowley and coworkers modified ligand **1.1b** by either exchanging the pyridine with quinoline or isoquinoline donors (**1.19a** and **1.19b**)⁸³ or by equipping the pyridines with amino groups at *ortho*- or *meta*-position to the nitrogens (**1.19c** and **1.19d**).⁹⁷ Ligands **1.19b** and **1.19d** form, similarly to parent ligand **1.1b**, a lantern-shaped cage. However, self-assembly of **1.19a** and **1.19c** which possess modifications in closer proximity to the coordinating nitrogen resulted in the formation of a mesocate $\text{Pd}_2(\mathbf{1.19a})_4$, **1.20** and a helicate $\text{Pd}_2(\mathbf{1.19c})_4$ (Fig. 2c). For $\text{Pd}_2(\mathbf{1.19c})_4$, this was explained by hydrogen bonds between the amino groups in the twisted coordination environment.

2.2. Unusual dinuclear assemblies

Severin and coworkers observed that acridone-based helicate $\text{Pd}_2(\mathbf{1.16b})_4$, initially reported by us,⁸⁰ interacts with a fifth ligand **1.16b**, when Pd(II) and **1.16b** are combined in a 2 : 5 ratio. In order to get insight into the nature of this association, the authors added **1.16c** which is a derivative of **1.16b**, lacking the nitrogen donors. The authors observed the formation of a 1 : 1 complex between **1.16c** and helicate $\text{Pd}_2(\mathbf{1.16b})_4$. The association of compound **1.16c** with $\text{Pd}_2(\mathbf{1.16b})_4$ is presumably driven by strong ligand-to-ligand π - π stacking interactions. For obtaining a discrete penta-stranded helicate Pd_2L_5 , the authors exploited the multiple inward pointing carbonyl groups of the helicate: oxophilic, high-valency La(III) binds to the pocket of the helicate and stabilizes the arrangement of five ligands **1.16b** *via*



coordination bonds. In $\text{La(III)}@Pd_2(\mathbf{1.16b})_5$, **1.21**, three ligands coordinate to both Pd(II) centres while the two remaining ones coordinate *via* one donor group only, giving rise to an assembly of overall low symmetry (Fig. 3a).⁸⁴

Our group obtained a complex, interlocked architecture $Pd_2(\mathbf{1.22})_4$, **1.23** with bis-pyridyl ligand **1.22** which is based on a peptidic macrocycle (Fig. 3b). The assembly is chiral and represents an entirely new catenated motif for Pd(II) architectures. The high ligand flexibility allows it to chelate to a single metal ion to give mononuclear complex $Pd(\mathbf{1.22})_2$, possessing a butterfly structure. Self-assembly in acetonitrile, followed by the addition of chloride or triflimide salts, however, eventually led to the formation of a dinuclear complex. The Pd(II) centres in the C_2 -symmetric structure are chelated in a *cis*-configuration and the two $Pd(\mathbf{1.22})_2$ moieties, also termed lemniscates, are held together by two mechanical bonds. A closer look at the X-ray structure reveals that this peculiar architecture is stabilized by (a) a counter anion that binds between the two Pd(II) centres, (b) inter-ligand π - π stacking interactions and (c) embedding of

pyridine donor moieties in the macrocyclic backbone cavity of the adjacent ligand.⁹⁸

Moreover, we showed that the introduction of steric congestion on the donor site can result in the formation of Pd_2L_m , $m < 4$ species, thereby challenging the principle of maximum site occupancy. While pyridine-equipped ligand **1.24a** based on a dibenzo-2.2.2-bicyclooctane backbone assembles to form capsule $Pd_2(\mathbf{1.24a})_4$, the same ligand with 6-quinoline donors **1.24b** yields bowl $Pd_2(\mathbf{1.24b})_3(\text{CH}_3\text{CN})_2$, **1.25** when Pd(II) and **1.24b** are combined in a 2 : 3 ratio (Fig. 3c). This outcome can be explained by steric repulsion: in capsule $Pd_2(\mathbf{1.24b})_4$, the distances between protons ⁸H (red in Fig. 3c) of neighbouring ligands are smaller than the sum of their van der Waals distances, leading to steric repulsion, while they have more space and thus surmount this distance in bowl **1.25**.⁷⁶ Interestingly, when ligand **1.24c**, equipped with 3-quinoline donors is employed, capsule $Pd_2(\mathbf{1.24c})_4$ is formed as a thermodynamic product. For overcoming steric stress, the donor groups arrange in a propeller geometry, leading to an overall helical twist of the assembly. Acridine as a donor moiety introduces two hydrogen substituents that are oriented in a similar direction to the nitrogen donor vectors. Ligand **1.24d** hence has a further increased steric hindrance compared to **1.24b** and **1.24c**, favouring the formation of ring $Pd_2(\mathbf{1.24d})_2(\text{CH}_3\text{CN})_4(\text{Cl})_4$, **1.26**.⁸² This topological variance has consequences for the fullerene binding capabilities of the assemblies. While capsule $Pd_2(\mathbf{1.24a})_4$ is able to bind C_{60} but not C_{70} , the open bowl **1.25** allows for binding of both. Helically twisted $Pd_2(\mathbf{1.24c})_4$ and ring **1.26**, on the other hand, showed no or weak affinity towards fullerenes. Partial solvent exposure of the surface of C_{60} in bowl **1.25** was exploited for controlling the covalent functionalization of fullerenes with the bowl serving as a non-covalent protecting group.

Formation of Pd_2L_3 -type bowls along with Pd_2L_2 -type rings was also observed for bis-picolyl ligands, *i.e.* when the steric congestion at the pyridine is increased through the introduction of methyl groups.⁹⁹ Furthermore, coordination rings are also formed when doubly protected Pd(II) cations or a moiety such as $PdCl_2$ is employed.^{100–102}

Recently, we reported a self-penetrated homoleptic assembly based on ligand **1.27** which possesses two key attributes: (a) strongly convergent donor vectors and (b) diketopyrrolopyrrole (DPP) backbones, known for their high propensity to interact *via* π - π stacking. Doubly-bridged ravel $Pd_2(\mathbf{1.27})_4$, **1.28** is obtained as a racemic mixture, belonging to point group D_2 . The two inner ligands adopt an S-shaped conformation which allows them to cross the centre of the assembly. The other two ligands sit on the outer positions and are helically twisted (Fig. 3d). Overall, four DPP moieties are stacked whereby neighbouring backbones are rotated by roughly 90°. The assembly is reminiscent of heteroleptic *trans*- $Pd_2A_2B_2$ cages,¹⁰³ which will be described later. In ligand **1.29** the DPP backbone is equipped with sterically demanding isoquinoline donors which led to the formation of self-penetrated bowl $Pd_2(\mathbf{1.29})_3$, **1.30**. In contrast to previously reported bowl structures, the central and one of the outer ligands adopt an S-shaped conformation. This allows three DPP moieties to stack and

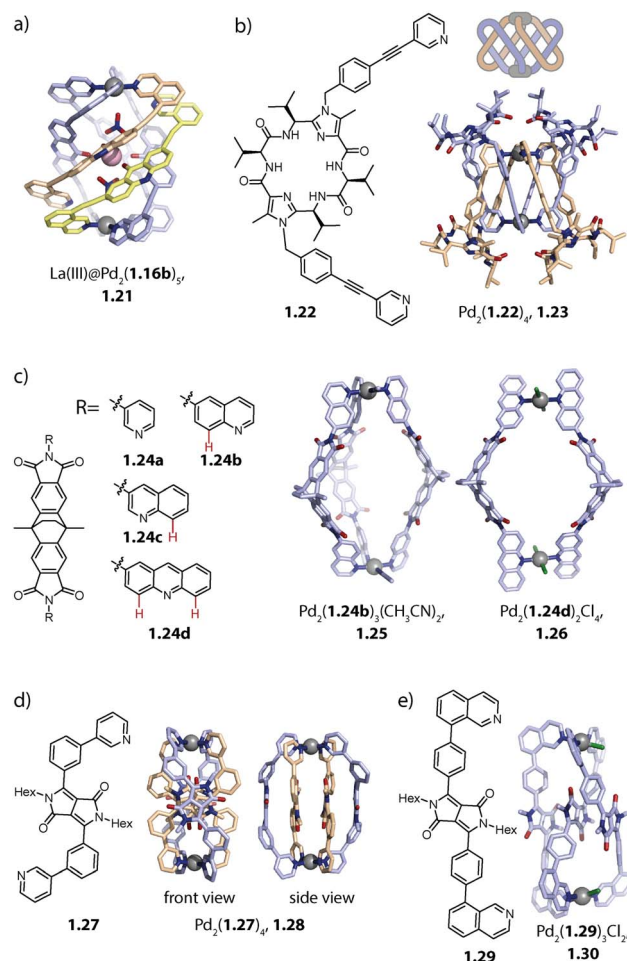


Fig. 3 (a) Penta-stranded helicate **1.21** templated by La(III) . (b) Doubly interlocked lemniscates **1.23**. (c) Bowl **1.25** and ring **1.26** (protons marked red would cause steric clash in a potential Pd_2L_4 arrangement). (d) Double-bridged ravel **1.28**. (e) Single-bridged ravel **1.30**. Positions on Pd(II) centres not saturated by N-heterocyclic ligands are either occupied by solvent or halide ligands.



gives rise to a C_2 -symmetric assembly, not leaving any cavity for guest binding (Fig. 3e).¹⁰⁴

In summary, here we showcased a series of experimentally observed topologies for assemblies possessing only two Pd(II) ions. The design principles for lantern-shaped cages and helicates are well established. For the former, collinear bis-monodentate ligands are utilized; helical twisting is driven by steric constraints, inter-ligand interactions, and/or strongly convergent donor vectors. We also touched upon examples where the steric demand of the donor groups resulted in the formation of more open bowl or ring structures. It is notoriously more difficult to design intricate self-penetrated or interwoven assemblies whose formation is often governed by close-ligand interactions, templation, solvent effects or a combination thereof.

3. Homoleptic multinuclear assemblies ($n = 3-8$)

Dinuclear cages (Chapter 2) and spheres Pd_nL_{2n} $n \geq 12$ have garnered great attention and for the latter, guidelines for their rational design have been described by Fujita and coworkers.¹⁰⁵ Furthermore, most functional assemblies reported so far belong to either of these classes.^{64,105} Less attention has been paid to assemblies with intermediate sizes ($n = 3-8$), which will be described in this chapter. We discuss the assemblies in the order of increasing nuclearity and complexity, starting with three- to five-membered rings. Next, we will discuss pseudo-tetrahedral and hexanuclear species such as cubes, octahedra,

and rings. Finally, we address more complex supramolecular coordination assemblies, such as interlocked cages. For clarity, different assemblies based on the same ligand are usually discussed together; hence, the order is not strictly followed.

3.1. Coordination rings

For the nuclearities $n = 3, 4$, and 5, the assemblies of the highest symmetry are rings. Fujita and coworkers synthesized the first three-membered ring in 2006 which will be described later.¹⁰⁶ The same group showed that ligand **2.1**, possessing a binding angle of 63° , assembles to form the entropically favoured three-membered ring $Pd_3(2.1)_6$, **2.2** in acetonitrile, while a four-membered ring $Pd_4(2.1)_8$, **2.3** is formed in DMSO (Fig. 4a). Interestingly, X-ray crystal structure analysis revealed that the cavity of **2.3** in the solid state is occupied by an aggregate of DMSO and nitrate molecules. Hence, the solvent dependency might be explained by a templation effect. Self-assembly in the presence of other anions yields mixtures of tri- and tetranuclear species, supporting this conclusion.¹⁰⁷ Previously, a four-membered ring Pd_4L_8 had only been reported by Chand, Fujita and coworkers in 2003.¹⁰⁸ A similar solvent-dependent equilibrium between 3- and 4-membered rings was recently described in mechanistic detail by Li and coworkers.¹⁰⁹

Previously, the same group also investigated the self-assembly of metalloligands **2.4a** and **2.4b** with binding angles of 60° , differing solely in their length and hence in their flexibility. While short ligand **2.4a** cleanly assembles to form four-membered ring $Pd_4(2.4a)_6$, **2.5**, ligand **2.4b** forms trinuclear ring $Pd_3(2.4b)_6$, **2.6** (Fig. 4b). Hence, the increased flexibility of

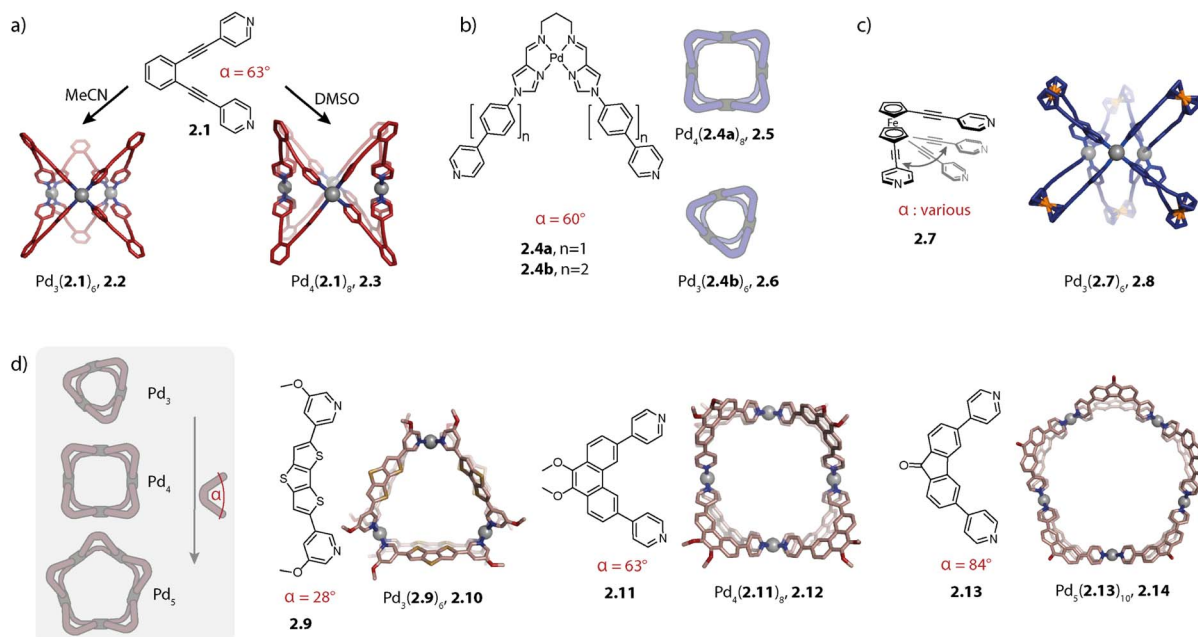


Fig. 4 Self-assembly of homoleptic rings Pd_nL_{2n} ($n = 3-5$). (a) Solvent-controlled self-assembly of three- and four-membered rings **2.2** and **2.3**. (b) Self-assembly of three- and four-membered rings **2.5** and **2.6** controlled by the ligand length. (c) Three-membered ring **2.8** based on a rotationally flexible ferrocene-based ligand. (d) Control over the ring size achieved through ligand angle adjustment (all X-ray structures; for **2.10**, a DFT model is shown). For Pd_6L_{12} rings, see Fig. 5b and 8d (α refers to the binding angle of the free ligand and the values for ligands **2.1**, **2.4**, **2.9**, **2.11**, and **2.13** were taken from ref. 112).



ligand **2.4b** allows the formation of the entropically favoured smaller assembly.¹¹⁰

Crowley and coworkers studied the self-assembly of ferrocene-based ligand **2.7**, which can, in contrast to the rather rigid ligands discussed beforehand, adopt a variety of different binding angles owing to its rotational flexibility (Fig. 4c). Interestingly, the three-membered ring Pd₃(**2.7**)₆, **2.8** was obtained as the sole species. Presumably, this is the smallest assembly accessible without significant ring strain; hence, the outcome can be rationalized based on entropic considerations.¹¹¹

We recently studied the self-assembly of a series of ligands under systematic variation of the binding angle α under similar conditions (DMSO, [Pd(CH₃CN)₄](BF₄)₂) for deriving a relationship between the ligand bent angle and ring nuclearity (Fig. 4d). Ligand **2.9** with an angle of 28° assembles to form the three-membered ring Pd₃(**2.9**)₆, **2.10** and phenanthrene-based ligand **2.11** with an angle of 63° forms, as was shown previously,⁸⁰ the four-membered ring Pd₄(**2.11**)₈, **2.12**. The binding angle of fluorenone based ligand **2.13** is further increased to 84° which allowed us to cleanly assemble the first non-templated five-membered ring Pd₅(**2.13**)₁₀, **2.14**. Self-assembly of a dibenzothiophene-based ligand with an intermediate binding angle of 78° leads to the formation of a mixture of four- and five-membered rings, showcasing that the ring size can be controlled by the bent angle when self-assembly conditions and ligand flexibility remain constant. It has to be noted that the self-assembly of the discussed ligands displayed solvent dependency and also tetra- and hexanuclear species, most probably octahedra, could be observed by mass spectrometric analysis. By treating the ligands as rigid triangles, a geometrical model was derived which reproduced the observed trends in a qualitative fashion. Major bottlenecks of such simple geometrical considerations, however, are omission of the ligand

flexibility and solvent effects, hence serving only as a rule of thumb approach.¹¹²

Crowley and coworkers synthesized a heterometallic Pd₃(PtL₂)₆ ring by employing a ligand with a bidentate and two monodentate binding sites. In the first step, Pt-based metal-oligand L₂Pt was synthesized through coordination *via* the bidentate site. This was followed by self-assembly with Pd(II). When L was reacted with only Pd(II) cations, a mixture of different assemblies was obtained and Pd₃(PdL₂)₆ was identified as the main component. The heterometallic ring was shown to incorporate planar aromatic guests within its clefts and was used for catalysing the [4 + 2] cycloaddition of anthracene with singlet oxygen.¹¹³

In 2014 and 2015, the groups of Mukherjee and Chand assembled a three-membered ring Pd₃L₆ from simple 1,3-diimidazole-benzene ligands.^{114,115} Chand showed that a similar ligand with benzimidazole donors also assembles to form a three-membered ring and that this assembly allows for gel formation owing to the additional π -surfaces.¹¹⁵

Furthermore, we equipped a photoswitchable dithienylethene (DTE) backbone with *para*-pyridine donors. Opening and closing of the DTE moiety upon irradiation with UV or visible light, respectively, is accompanied by a drastic change in the ligand bent angle. This allowed for reversibly switching between a mixture of rings (Pd₃L₆ : Pd₄L₈ 3 : 1) and a Pd₂₄L₄₈ sphere. While assembly in the presence of tetrafluoroborate anions afforded the aforementioned mixture of three- and four-membered rings, the trinuclear species was formed as the sole species with a Pd(II) nitrate salt. Presumably, this can be attributed to a templation effect.¹¹⁶

Templation of Pd₃L₆ by nitrate anions was also reported by Jung and coworkers. The authors showed that self-assembly of a very flexible bis-pyridyl ligand **2.15** with different Pd(II) precursors results in entropically favoured lemniscates Pd(**2.15**)₂, **2.16**. In contrast, when Pd(II) nitrate was used as the

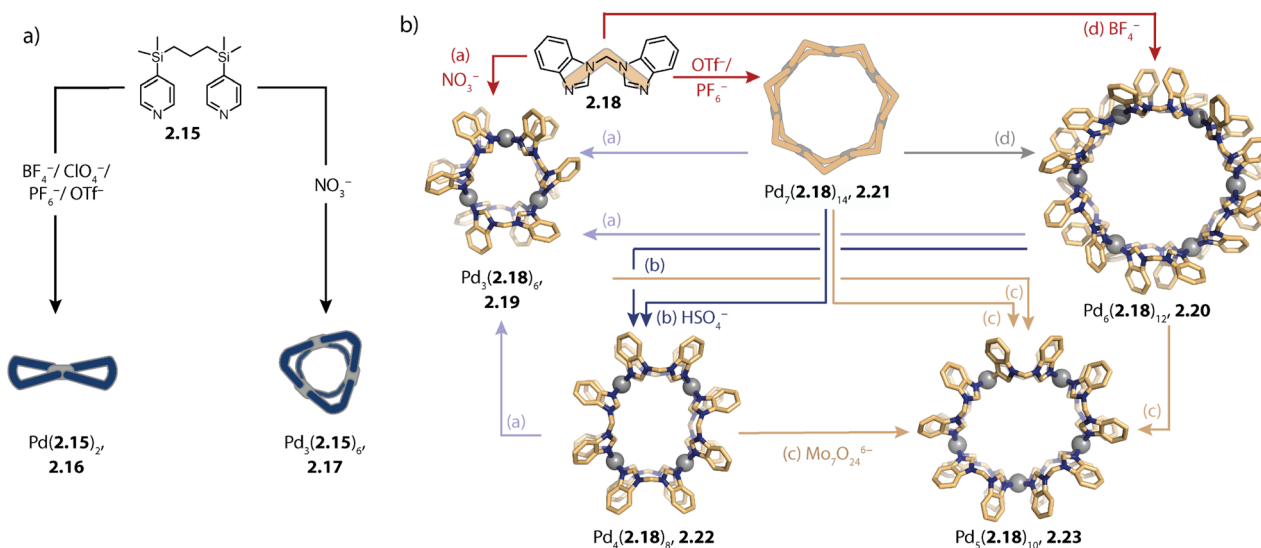


Fig. 5 Self-assembly of homoleptic rings Pd_nL_{2n} (n = 3–7) by templation. (a) Flexible ligand **2.15** self-assembles to form entropically favoured lemniscate **2.16** or *via* templation to the three-membered ring **2.17**. (b) Flexible, small ligand **2.18** assembles to form rings **2.19**–**2.23** of various sizes depending on the counter anion(s) present.



metal source, a clean three-membered ring $\text{Pd}_3(\mathbf{2.15})_6$, $\mathbf{2.17}$ was formed (Fig. 5a).¹¹⁷

In a similar fashion, Sun and coworkers showed nitrate templated formation of the three-membered ring $\text{Pd}_3(\mathbf{2.18})_6$, $\mathbf{2.19}$ (Fig. 5b). Surprisingly, rings of various nuclearities, ranging from $n = 3-7$ were assembled with ligand $\mathbf{2.18}$, strictly depending on the anion(s) present in solution. Akin to the three-membered ring, the six- and seven-membered rings $\text{Pd}_6(\mathbf{2.18})_{12}$, $\mathbf{2.20}$ and $\text{Pd}_7(\mathbf{2.18})_{14}$, $\mathbf{2.21}$ were obtained through direct assembly with the BF_4^- , or $\text{PF}_6^-/\text{OTf}^-$ salts, respectively (red arrows in Fig. 5b). In contrast, four-membered ring $\text{Pd}_4(\mathbf{2.18})_8$, $\mathbf{2.22}$ could only be accessed *via* ring-to-ring transformation by addition of HSO_4^- to preformed $\mathbf{2.20}$ or $\mathbf{2.21}$ (dark blue arrows) and five-membered ring $\text{Pd}_5(\mathbf{2.18})_{10}$, $\mathbf{2.23}$ by addition of $\text{Mo}_7\text{O}_{24}^{6-}$ to any of the preformed rings (beige arrows). From the different transformation processes, the authors concluded that the templation effect for this system is decreased in the order $\text{Mo}_7\text{O}_{24}^{6-} > \text{NO}_3^- > \text{SO}_4^{2-} > \text{BF}_4^- > \text{PF}_6^- = \text{OTf}^-$. The high adaptivity was attributed to (a) the methylene group which enables rotation as well as bending and hence adoption of numerous binding angles and (b) the acidic CH protons which might contribute to specific anion binding.¹¹⁸

3.2. Tetrahedra

In contrast to *para*-pyridine-based ligands which often form three-membered and/or four-membered rings,^{107,110,112} small ligands equipped with *meta*-pyridine donors tend to assemble to form pseudo-tetrahedra as tetranuclear species.^{106,119,120} The first example of a tetrahedron Pd_4L_8 was reported by Fujita and coworkers in 2006. The authors showed that short ligand $\mathbf{2.24}$

forms a triangular species $\text{Pd}_3(\mathbf{2.24})_6$, $\mathbf{2.25}$ with triflate as the counter anion while a tetrahedron $\text{Pd}_4(\mathbf{2.24})_8$, $\mathbf{2.26}$ was obtained when a nitrate salt was employed (Fig. 6a). Note that Pd_4L_8 assemblies can only give rise to pseudo-tetrahedral structures in which two oppositely arranged edges are doubly bridged while the remaining four edges are bridged by a single ligand each. Hence, each vertex consists of a tetra-coordinated $\text{Pd}(\text{II})$ cation (whereas in a real tetrahedron, each vertex connects only three edges). Elongation of the ligand with one phenyl linker results in a dynamic equilibrium between both species: with nitrate, the triangular ring and the tetrahedron coexist in a ratio of 1 : 1 while the ring is favoured when a $\text{Pd}(\text{II})$ triflate salt is employed. With tosylate, the tetrahedral species is yielded as the sole assembly.¹⁰⁶

Mukherjee and coworkers studied the self-assembly of small bis-*meta*-pyridyl ligands of a similar geometry. Self-assembly of ligand $\mathbf{2.27}$ with a $\text{Pd}(\text{II})$ nitrate salt yielded, similar to $\mathbf{2.24}$, tetrahedral cage $\text{Pd}_4(\mathbf{2.27})_8$, $\mathbf{2.29}$ exclusively. In contrast, ligand $\mathbf{2.30}$ assembles to form triangular $\text{Pd}_3(\mathbf{2.30})_6$, $\mathbf{2.31}$ under the same conditions (Fig. 6b). The self-assembly of ligand $\mathbf{2.27}$ was furthermore studied with $\text{Pd}(\text{II})$ tetrafluoroborate salt in DMSO and ACN. While the tetrahedral species $\mathbf{2.29}$ was favoured in ACN with this counter anion as well, a mixture of ring $\mathbf{2.28}$ and tetrahedron $\mathbf{2.29}$ was obtained in DMSO. Self-assembly of ligand $\mathbf{2.30}$ yielded a dynamic mixture in both solvents with the triangular species $\mathbf{2.31}$ as the main component in ACN and tetrahedral $\mathbf{2.32}$ in DMSO. From the investigation of the equilibrium of both DMSO systems at various temperatures, the authors were able to conclude that the tetrahedra $\mathbf{2.29}$ and $\mathbf{2.32}$ are entropically favoured over the corresponding rings but are

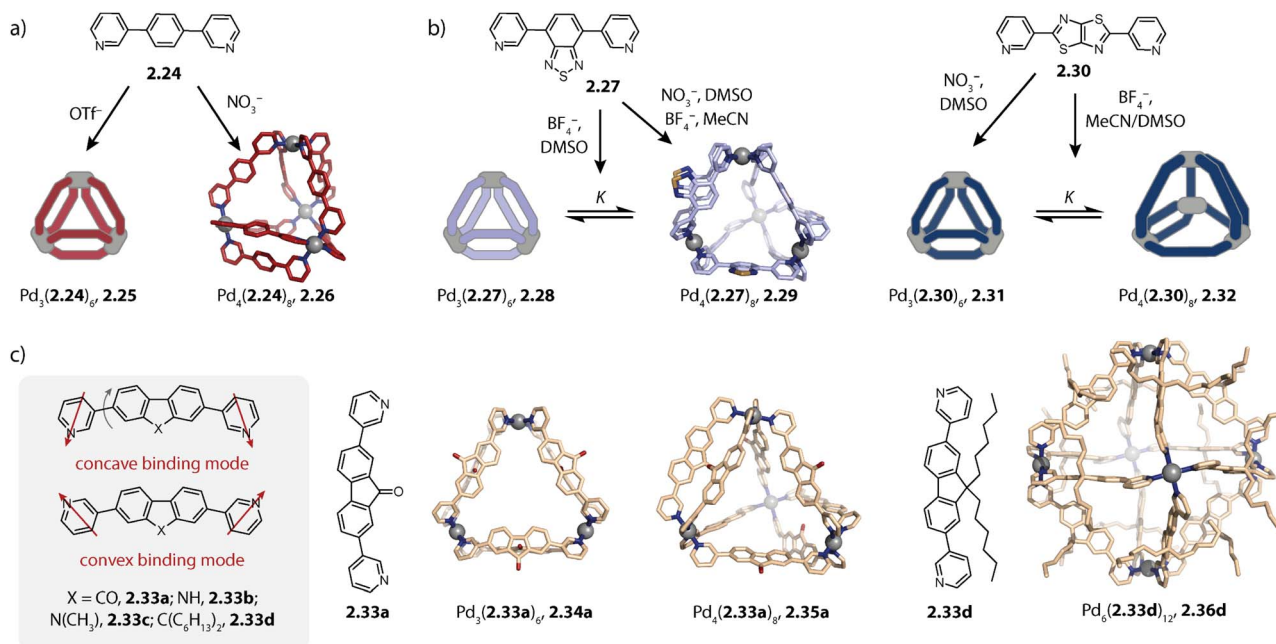


Fig. 6 Self-assembly of Pd_3L_6 rings and Pd_4L_8 pseudo-tetrahedra with bis-*meta*-pyridyl ligands. (a) Counter-anion dependent self-assembly of $\mathbf{2.25}$ and $\mathbf{2.26}$. (b) Self-assembly of three-membered rings $\mathbf{2.28}$ or $\mathbf{2.31}$ and tetrahedra $\mathbf{2.29}$ or $\mathbf{2.32}$, dependent on the counter anion and solvent. (c) Fluorene-derived ligands $\mathbf{2.33}$ can adopt different binding modes, allowing for assembly to Pd_3L_6 rings, Pd_4L_8 tetrahedra, and Pd_6L_{12} octahedra.



enthalpically disfavoured; hence, the equilibrium shifts towards the larger species at higher temperatures.¹²⁰

Moreover, our lab investigated the self-assembly of fluorene-derived ligands **2.33**.^{119,121} Considering the flat ligands, two binding situations, namely the concave and the convex mode, can be differentiated (Fig. 6c). A binding mode in which the two donor vectors are oriented in opposite directions would most probably not lead to a discrete assembly. Fluorenone-¹¹⁹ and *N*-methyl carbazole¹²¹ based ligands **2.33a** and **2.33c** assemble to form a mixture of the three-membered ring $\text{Pd}_3(\mathbf{2.33a/c})_6$, **2.34a/c** and tetrahedral $\text{Pd}_4(\mathbf{2.33a/c})_8$, **2.35a/c** in acetonitrile. In contrast, ligand **2.33b** based on *N*-unsubstituted carbazole cleanly assembles to form the tetrahedral species **2.35b**. Even though the ligands in the assembly are not flat, the two distinct binding modes are observed in the X-ray crystal structure of **2.35a** (Fig. 6c, middle): the ligands sitting on the singly bridged edges adopt a rather convex binding mode, and hence the backbone functionality points outside, while the ligands occupying the doubly bridged edges adopt a concave binding mode. When fluorene-based ligand **2.33d** with two hexyl chains is employed, octahedral $\text{Pd}_6(\mathbf{2.33d})_{12}$, **2.36d** is obtained. All ligands adopt a convex binding mode, allowing for assembly to this higher nuclearity species. Occupation of double-bridged edges is not possible with **2.33d** owing to steric hindrance; hence, formation of homoleptic rings and tetrahedra is impeded.¹¹⁹ It is worth noting that the backbone steric bulk does serve as a topology-controlling moiety here that does not compromise the size of the inner cavity. This contrasts with the first study on the backbone steric bulk-controlled formation of octahedra by Severin and coworkers where the overall size of the ligand backbone was increased so much that it reaches into the central cavity, as will be discussed later.^{122,123}

Lützen and coworkers applied chiral BINOL-based ligand **2.37** for assembling a tetranuclear species $\text{Pd}_4(\mathbf{2.37})_8$, **2.38** in which the Pd(II) ions are arranged in a tetrahedral fashion (Fig. 7a). In contrast to the aforementioned tetrahedra possessing singly and doubly bridged edges, all neighboring metal ions are doubly bridged in **2.38**. In this distorted macrocyclic structure, the ligands adopt two distinct conformations: ligands bridging the long edges have a W-shaped conformation with rather divergent binding angles while ligands bridging the short edges possess a C-shaped conformation with a smaller binding angle. This arrangement results in three pockets: two smaller peripheral pockets, each encapsulating one tetrafluoroborate anion, and a central cavity. Self-assembly of **2.37** with other Pd(II) sources did not result in a defined species, showing that the tetrafluoroborate anions serve as the template for **2.38**.¹²⁴

We recently reported an unprecedented structural switch between a lantern-shaped cage $\text{Pd}_2(\mathbf{2.39})_4$, **2.40** and a distorted tetrahedron $\text{Pd}_4(\mathbf{2.39})_8$, **2.41** (Fig. 7b). Bis-pyridyl ligand **2.39** with an azulene backbone assembles with Pd(II) to a classical lantern-shaped cage **2.40**, as can be expected based on the geometry of the ligand (Chapter 2.1). Conversion into **2.41** was achieved upon addition of 1,4-benzene-bisulfonate (BEN). While the conformation of the ligands sitting on the doubly bridged edges closely resembles the situation in the lantern-shaped cage, the pyridines of the ligands on the singly

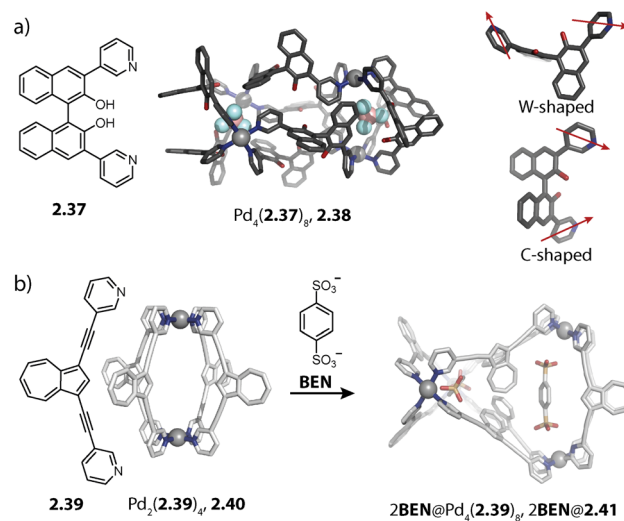


Fig. 7 Templated self-assembly of the Pd₄L₈ distorted macrocycle and tetrahedron. (a) Chiral ligand **2.37** forms distorted tetrahedron **2.38** in which the ligands are present in a W-shaped and a C-shaped conformation. Two tetrafluoroborate counter anions are encapsulated in the outer pockets. (b) Azulene-based ligand **2.39** assembles to form lantern shaped cage **2.40** which shows guest-induced conversion to a distorted tetrahedron **2.41**.

bridged edges are flipped by approximately 180°. The shorter Pd–Pd distances, connecting the nodes of the double-bridged edges, match the size of the BEN guest, allowing for close attractive electrostatic interactions. Hence, maximizing host-guest interactions is presumably the driving force for the observed structural switch, aided by an energetic contribution resulting from a circular arrangement of the dipolar backbones of the singly bridged edges. Intriguingly, the switch could be reversed by addition of a competitive host.⁷⁹

3.3. Hexanuclear species

Rigid ligands with nearly rectangular geometry assemble to form cubes in which the Pd(II) ions are arranged in an octahedral fashion. This can be regarded as an extension of the above discussed series of rigid ligands to even larger binding angles. The first cube $\text{Pd}_6(\mathbf{2.41b})_{12}$, **2.43** was reported by Fujita and coworkers in 2009 and is assembled from dibenzofurane ligand **2.41b** (Fig. 8a).¹²⁵ Also, smaller ligand **2.41a** and carbazole-based ligand **2.42** assemble to form analogous cubes, showcasing the reliability of the design approach based on the binding angle of rigid ligands. Reek and coworkers showed that cubes $\text{Pd}_6(\mathbf{2.41a})_{12}$ and $\text{Pd}_6(\mathbf{2.42})_{12}$ bind fullerene and that the binding affinity can be tuned by altering the electronic properties of the ligands.¹²⁶

Furthermore, the group of Lützen reported BINOL-based ligands which possess, in contrast to ligand **2.37** discussed above, *para*-pyridine donor groups and hence larger binding angles. The ligands self-assemble to form Pd₆L₁₂ cubes and Pd₁₂L₂₄ spheres.¹²⁷

Mukherjee and coworkers reported the first octahedra based on bis-imidazole ligands **2.44** (Fig. 8b). Similar to the cube-like



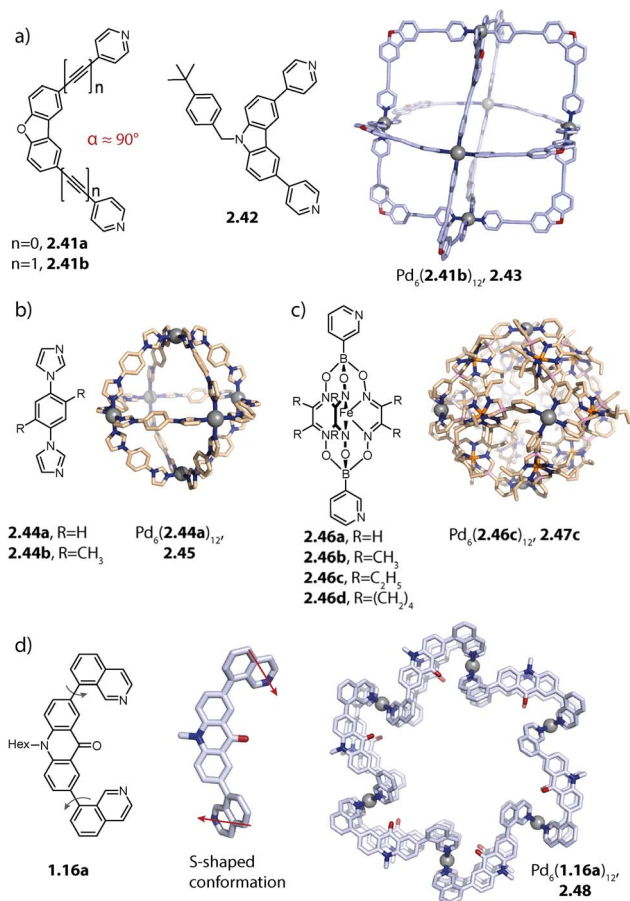


Fig. 8 Homoleptic Pd₆L₁₂ assemblies. (a) Cubes (with an octahedral metal arrangement) based on rigid bis-*para* pyridyl ligands with binding angles close to 90°. (b) Octahedron **2.45** based on a short bis-imidazole terminated ligand. (c) Octahedron **2.47c** based on bis-*meta* pyridyl ligands with a sterically crowded backbone. (d) Six-jagged ring **2.48** based on flexible ligand **1.16a** adopting an S-shaped conformation.

structure, the Pd(II) ions in Pd₆(**2.44**)₁₂, **2.45** are arranged in an octahedral fashion but due to the different ligand geometry, neighboring Pd(II) are bridged in a rather linear fashion.¹²⁸

Severin and coworkers reported clathrochelate ligands **2.46** which possess a similar geometry compared to ligand **2.24** introduced by the group of Fujita. However, the increased thickness of the backbone allowed for topology control: while less bulky **2.46a** assembles to form the entropically favoured tetrahedron Pd₄(**2.46a**)₈, ligands **2.46b–d** with increased steric bulk only allow the formation of larger octahedra Pd₆(**2.46b–d**)₁₂, **2.47b–d** (Fig. 8c). Presumably, the larger species is favoured here in order to avoid repulsive steric interactions between the ligands in the smaller assembly. Furthermore, the authors showed with the help of enforced disassembly experiments that the decreased thickness of ligand **2.46d** as compared to **2.46c** translates into an enhanced thermodynamic stability of **2.47d**. In turn, narcissistic self-sorting was observed when ligand **2.46d** was combined with ligand **2.24**. In contrast, a combination of **2.46c**, **2.24**, and Pd(II) resulted in a library of mixed-ligand species.^{122,123}

Recently, we reported the surprising conformational flexibility of short acridone-based ligand **1.16a**. In solution, helically twisted Pd₂(**1.16a**)₄ is formed which is, depending on the solvent and counter anion, transformable into its mesocate isomer. Surprisingly, when we attempted to crystallize the mesocate, six-jagged ring Pd₆(**1.16a**)₁₂, **2.48** was obtained (Fig. 8d). Here, the ligands adopt an S-shaped conformation and pairs of ligands bridge neighbouring Pd(II) ions, resulting in an overall C_{6h}-symmetric structure. Crystal packing analysis revealed that the rings stack cofacially to form columnar, hexagonal channels that are separated by triangular gaps. The templated six-membered ring **2.22** reported by Sun and coworkers as described earlier and this solid-state structure are the first reported Pd₆L₁₂ rings.⁹⁴

3.4. Interlocked and other unusual motifs

Lantern-shaped coordination cages can undergo dimerization through catenation. Thereby, the upper part of the ligands of one cage penetrates through the windows of the other cage. The most frequently observed interlocked double cages possess a fourfold symmetry and an alternating stack of Pd(II) cations and counter anions which reduces the Coulomb repulsion between the closely positioned Pd(II) ions. We have reviewed interpenetrated double cages in depth in the past¹²⁹ and here, we will focus on some factors that allow for controlling interpenetration and on a recently reported new motif.

In 2008, Kuroda and coworkers reported Pd₄L₈ interlocked double cages for the first time. In this work, a combination of an organic ligand and Pd(NO₃)₂ in DMSO led to the formation of a monomeric cage Pd₂L₄ as a kinetic intermediate which was then converted into the quadruply interlocked double cage Pd₄L₈. Interestingly, when other Pd(II) sources were used for the self-assembly, the equilibrium was shifted towards the monomeric cage. Also, the addition of a naphthalene-monosulfonate guest led to a preference for the monomeric cage.¹³⁰

Later, we reported that dibenzosuberone-based ligand **2.48** assembles to form a similar interpenetrated cage Pd₄(**2.48**)₈, **2.49** with [Pd(CH₃CN)₄](BF₄)₂ in acetonitrile (Fig. 9b). Akin to the report of Kuroda and coworkers, the monomeric cage was observed to be a kinetic intermediate. In contrast to ligand **2.48** with alkyne spacers, double cage formation is not observed for ligand **2.50** with a shorter N⋯N distance (11.36 Å vs. 16.34 Å) under the same conditions. The pockets of a hypothetical double cage Pd₄(**2.50**)₈ would be too small for the incorporation of BF₄⁻ (or any other) anions, resulting in a preference for the monomeric cage Pd₂(**2.50**)₄.¹³¹

For further controlling double- vs. monomeric cage formation, we investigated how structural modifications of the ligand **2.48** affect its ability to dimerize. Elongation of the linker by exchanging the alkyne with a phenyl spacer (**2.51**) and desaturation of the ligand backbone (**2.52a**) as well as modifications of the pyridines of **2.52** with methoxy ethoxy chains (**2.52b**) did not impede interpenetration. In contrast, deoxygenation of the backbone (**2.53**) led to formation of the monomeric cage Pd₂(**2.53**)₄. Close examination of the X-ray crystal structure in combination with DFT calculations revealed that the double



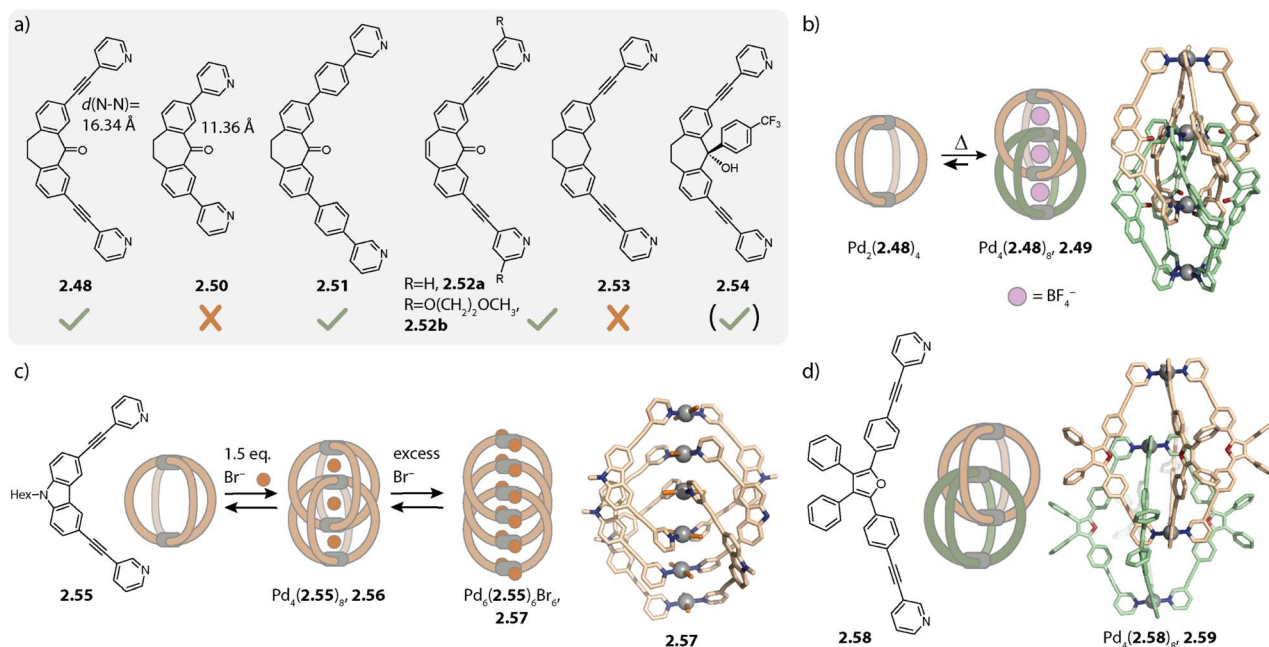


Fig. 9 Interlocked double cages Pd_4L_8 . (a) Structures of suberone-based ligands 2.48–2.54 (check mark: forms a double cage; cross: does not form a double cage). (b) Self-assembly of monomeric cage $\text{Pd}_2(\text{2.48a})_4$ as a kinetic product and double cage 2.49a as the thermodynamic product. (c) Halide induced formation of double cage 2.56 and ligand release upon addition of excess halides to form triple catenated 2.57 in the solid state. (d) Triply interlocked double cage 2.59 with furane based ligand 2.58.

cage structure 2.49 seems to be further stabilized by interactions between the oxygens of the carbonyl groups and the closely packed inner $\text{Pd}(\text{pyridine})_4$ plane of the other cage unit.¹³² This interaction, however, was not necessarily required for the interlocked cage reported by the group of Kuroda¹³⁰ and for double cages with other ligands based on phenothiazine¹³³ or helicene,¹³⁴ reported by us.

The propensity for intercalation is also altered when inward-pointing steric bulk is installed on the ligand backbone. Derivative 2.54 which possesses the same N...N distance and binding angle as compared to 2.48 forms lantern-shaped cage $\text{Pd}_2(\text{2.54})_4$ under similar self-assembly conditions to those used for obtaining double cage 2.49. However, the addition of chloride to $\text{Pd}_2(\text{2.54})_4$ again resulted in double cage formation. This can be rationalized based on steric considerations: due to the comparably large size of BF_4^- anions which would have to be encapsulated in the central pocket, the outer pockets of a potential cage $\text{BF}_4^-@(\text{Pd}_4(\text{2.54}))_8$ would be small, resulting in repulsive interactions between the attached aromatic groups. However, with chloride as the central anion, the outer pockets are enlarged, offering sufficient space for the bulky substituents.¹³⁵ Yamashina, Toyota and coworkers showed, on the other hand, that exohedral steric bulk can also control dimerization.¹³⁶

Similar chloride-induced interpenetration was observed with carbazole-based ligand 2.55 possessing shorter N...N distances as compared to 2.48. Only halides Cl^- and Br^- , but not BF_4^- , can be sandwiched between the $\text{Pd}(\text{py})_4$ planes, allowing for catenation of two cages to give $\text{Pd}_4(\text{2.55})_8$, 2.56 (Fig. 9c). In this case, addition of stoichiometric amounts of halides led,

however, to a mixture of the interlocked double cage with the monomeric cage and the free ligand. Addition of over-stoichiometric amounts of halides leads to further ligand release. This can be explained by competition between halides and organic ligands coordinating to $\text{Pd}(\text{II})$. Unexpectedly, triple catenated $(\text{PdBr}_2)_6(\text{2.55})_6$, 2.57, carrying no net charge, was crystallized. Here, the $\text{Pd}(\text{II})$ ions are arranged in a linear stack with noticeably shorter $\text{Pd}\cdots\text{Pd}$ distances as compared to the distances in the corresponding double cage (4.36 Å vs. 6.73 Å). The bromide ligands are arranged in a helical fashion and the close proximity of the ligand backbones suggests favourable π - π interactions to be the key driving force for the observed structure.¹³⁷

Bloch and coworkers reported the halide-triggered assembly of double cages that bind bisulfates with high affinity.¹³⁸ The group of Hiraoka described a system of transformable quadruply interpenetrated cages in which different halide-dependent states lead to different reactivities.¹³⁹

Most often, the degree of catenation is related to the symmetry of either the assembly or the ligand employed. This holds for the described quadruply interlocked cages with C_4 or D_4 symmetry, for triply interlocked C_3 symmetric cages of Fujita,¹⁴⁰ and for quintuply interlocked double cages based on C_5 symmetric ligands reported by Nitschke and coworkers.¹⁴¹ However, recently, we reported unprecedented triply interlocked double cage $\text{Pd}_4(\text{2.58})_8$, 2.59 with long furane-based ligand 2.58 (Fig. 9d). Similar to what has been reported for quadruply interlocked double cages, monomeric cage $\text{Pd}_2(\text{2.58})_4$ was formed as a kinetic product. Heating led to a mixture of the monomeric cage with quadruply and triply



interlocked double cage **2.59**. At room temperature, the latter was observed to be the thermodynamic product. Here, two opposing windows of a first cage are each penetrated by one bridging ligand of a second cage while a third window of the first cage is penetrated by two ligands of the second cage. This leaves one window of the first cage free. The peculiar structure is stabilized by π - π interactions between the electron rich furane backbones and the electron poor pyridines. A reference ligand with the same geometry as **2.58** but a silole backbone that possesses additional methyl substituents assembles to form a monomeric cage. Due to the methyl substituents, π - π interactions involving the backbones are hampered, which highlights the necessity of intense π - π interactions for the formation of an interlocked cage.¹⁴²

Aside from a simple three-membered ring, a double-trefoil knot structure $\text{Pd}_3(\mathbf{2.60})_6$, **2.61** was observed for the stoichiometry Pd_3L_6 when long concave ligand **2.60** was employed (Fig. 10a). Similar to the ring, each Pd(II) pair is bridged by two organic ligands; however, the ligands are intertwined. This allows for close interactions between ligands **2.60** which possess multiple sites for π - π and H-bond interactions. Among the different possibilities for the arrangement of the ligands in **2.61**, the peculiar structure was assigned employing a combination of NOESY NMR spectroscopy and molecular modelling, considering the symmetry of the ligands and the measured *vs.* calculated H...H distances. Double trefoil knot **2.61** consists of two chiral hemispheres with opposite configurations, forming an overall *meso* structure. The lack of preference for chiral

guests of opposite configurations is thus in full accordance with the proposed structure.¹⁴³

Severin and coworkers recently reported a structural switch from acridone based helicate $\text{Pd}_2(\mathbf{1.16b})_4$ to tetranuclear low symmetry assembly $\text{Pd}_4(\mathbf{1.16b})_8$, **2.62** upon addition of a Li salt. The structure consists of two $\text{Pd}_2(\mathbf{1.16b})_3$ distorted bowls that are bridged by two more ligands **1.16b** which sit in the centre of the structure and are involved in π - π interactions (Fig. 10b). Two binding pockets are each occupied by one LiBF_4 and a water molecule. The Li cation is stabilized by a tetrahedral environment of BF_4^- , water, and two carbonyl oxygens of the acridone ligands. The water molecule is furthermore engaged in H-bonds to two adjacent carbonyl oxygen atoms. Under strictly dry conditions, the peculiar, compact structure is not observed, underscoring the importance of the water molecules. Furthermore, the structural switch was selective for Li as it could not be induced by addition of other alkali metal salts.¹⁴⁴ The same group incorporated lantern-shaped cages with polymerizable appendices as dynamic crosslinks in acrylamide hydrogels.¹⁴⁵

Lützen and coworkers reported a cage-in-ring assembly $\text{Pd}_2(\mathbf{2.63})_4 @ \text{Pd}_4(\mathbf{2.63})_8$, **2.64**, resembling a rotaxane (Fig. 10d). The shape of ligand **2.63** reminds to simple banana-shaped ligands known to assemble to form lantern-shaped cages Pd_2L_4 . However, it stands out due to the length of its arms and possesses BODIPY and alkoxy substituents, having a propensity for engaging in π - π and van-der-Waals interactions, respectively. In the centrally positioned lantern-shaped cage $\text{Pd}_2(\mathbf{2.63})_4$, the ligands adopt a C-shaped conformation, leading

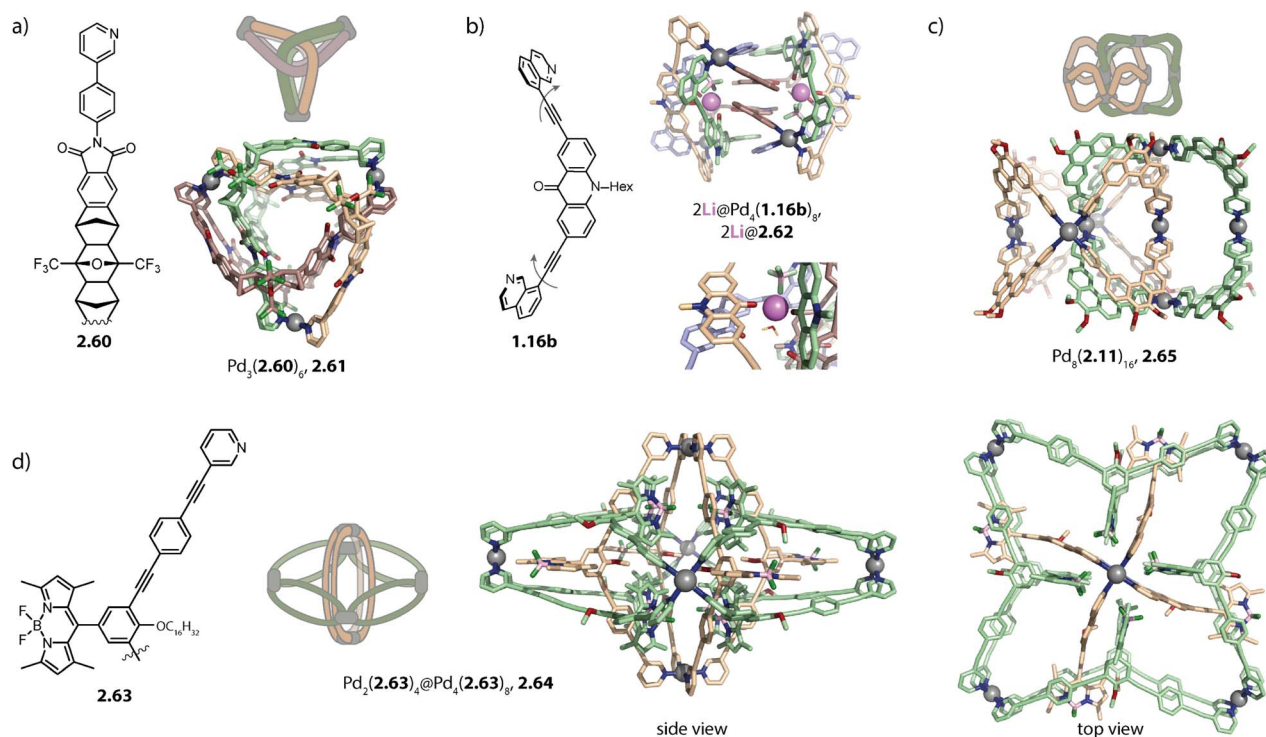


Fig. 10 Interwoven and interpenetrated structures Pd_nL_{2n} , $n = 3-8$. (a) Double trefoil knot **2.61**. (b) Li(i) templated low-symmetry assembly **2.62**. (c) Octanuclear assembly **2.65** consisting of two four-membered rings **2.11** that are interlocked and rotationally displaced by 90° . (d) Pseudo-rotaxane cage-in-ring assembly **2.64** formed by BODIPY-functionalized ligand **2.63**.



to outward-pointing BODIPY-moieties and inward-pointing alkoxy chains. The ligands forming the four-membered ring $\text{Pd}_4(\mathbf{2.63})_8$ possess a W-shaped conformation in which the pyridine moieties are flipped by approximately 180° . Hence, the BODIPY substituents are endohedrally located allowing for π - π interaction with the backbones of $\text{Pd}_2(\mathbf{2.63})_4$. Additionally, the outer ring possesses clefts which enable the intercalation of the BODIPY moieties of $\text{Pd}_2(\mathbf{2.63})_4$.¹⁴⁶

Catenation of Pd_nL_{2n} architectures with $n > 2$ was so far only observed for four-membered rings **2.12** based on phenanthrene ligand **2.11**. Octanuclear $\text{Pd}_8(\mathbf{2.11})_{16}$, **2.65** consists of two mechanically interlocked D_{4h} -symmetric rings **2.12** that are rotationally displaced by 90° (Fig. 10c). The assembly displays D_{2d} symmetry and can be described as a huge Hopf-link. Noteworthy, the interlocked motif was only observed in the presence of nitrate, either upon direct assembly with $\text{Pd}(\text{NO}_3)_2$ or upon addition of a nitrate salt to a preformed mixture of tri- and tetranuclear assemblies. Presumably, nitrate anions possess the ideal size for templating the dimerization. Furthermore, switching the solvent from acetonitrile to DMSO impeded interlocking, similar to the dimerization of most lantern-shaped cages. Furthermore, higher order assembly to vesicular structures was achieved by equipping ligand **2.11** with hexyloxy chains. Each modified **2.65** assembly carries 32 hexyloxy chains, promoting aggregation *via* hydrophobic interactions in acetonitrile.¹⁴⁷

To conclude, assemblies of higher nuclearity can be accessed with ligands possessing large bent angles, with flexible ligands and a suited template, or through interlocking of lower nuclearity assemblies. Starting off with architectures of rather high symmetry, namely rings, tetrahedra, and octahedra, we presented how the assembly size can be rationalized with the bent angle of the underlying ligand. We have also discussed more sophisticated architectures such as a cage-in-ring assembly, a double trefoil knot, and interlocked rings, among others, found serendipitously. Here, solvent and templation effects as well as ligand flexibility challenge the rationalization and prediction of the outcome.

4. Integrative self-sorting

As we can see, the investigation of homoleptic cage assembly has already produced an astonishing variety of topologies and it has enabled the rationalization of some key design principles. However, the functionality of such cages is inevitably limited as they carry only one kind of organic ligand – the building block that usually imbues the assembly with its function. For overcoming this constraint, multiple kinds of ligands can be combined, giving rise to so-called heteroleptic assemblies.^{64,148} The synthesis of such multicomponent cages is, however, not trivial in terms of the self-assembly strategy: when ligands of too similar size and shape are combined, the formation of an entropically favoured statistical mixture (*i.e.* assemblies with different stoichiometries and stereo configurations) is usually observed.^{148,149} The drawback of working with such mixtures is that only a subset of assemblies then comes with the desired functionality, and observed properties may not be correlated

with a particular structure, impeding the optimization of the most selective and capable derivatives. For limiting the number of possible stoichiometries and isomers, *cis*-protected $\text{Pd}(\text{II})$ sources have been successfully employed in the past. This has been combined with steric constraints by Fujita,¹⁵⁰ with differently charged donors by Mukherjee and others,^{151,152} as well as with variation in the ligand denticity¹⁵³ to achieve integrative self-sorting. Subsequently, strategies for the non-statistical self-assembly of bis-monodentate organic ligands carrying pyridine, isoquinoline, or imidazole donors with naked $\text{Pd}(\text{II})$ ions were developed by Crowley,¹⁵⁴ Hooley,¹⁵⁵ Severin,¹⁵⁶ and us.^{80,99,121,157} We will first describe Pd_2L_4 heteroleptic cages as they represent the most basic and largest family and will subsequently address heteroleptic cages of higher nuclearity.

4.1. $\text{Pd}_2\text{A}_2\text{B}_2$ cages

Inspired by the work of Fujita and coworkers on heteroleptic assemblies based on *cis*-protected $\text{Pd}(\text{II})$,¹⁵⁰ we exploited steric constraints at the donor sites for yielding $\text{Pd}_2\text{A}_2\text{B}_2$ -type cages with naked $\text{Pd}(\text{II})$ cations. For this, we coined the term “coordination sphere engineering” (CSE). Picolyl ligands **o-3.1** and **i-3.1** differ from the typical banana-shaped ligands discussed in Chapter 1 in the methyl-group adjacent to the nitrogen donor, either pointing outside (**o-3.1**) or inside (**i-3.1**) the tentative cavity of the respective assembly (Fig. 11a). Increased steric repulsion renders potential homoleptic species $\text{Pd}_2(\mathbf{o-3.1})_4$ or $\text{Pd}_2(\mathbf{i-3.1})_4$ energetically unfavourable; indeed, bowls or mixtures of bowls and rings are obtained when only one kind of ligand is employed. In contrast, the combination of both ligands **o-3.1** and **i-3.1** in a 1 : 1 ratio results in the formation of heteroleptic cage *cis*- $\text{Pd}_2(\mathbf{o-3.1})_2(\mathbf{i-3.1})_2$, **3.3**. Here, the repulsive steric interactions are decreased while the principle of maximum site occupancy is satisfied. DFT calculations and the X-ray crystal structure of a mononuclear $\text{Pd}(\text{picolyl})_4$ complex support that the *cis*-isomer possesses higher stability as compared to a tentative *trans*-isomer. The combination of acridone-based **o-3.1** with phenothiazine-based **i-3.2** led to a similar $\text{Pd}_2(\mathbf{o-3.1})_2(\mathbf{i-3.2})_2$ cage, *i.e.* the platform tolerates incorporation of two different functional backbones. Surprisingly, the reverse combination of **i-3.1** with **o-3.2** did not allow for clean self-assembly to heteroleptic $\text{Pd}_2(\mathbf{i-3.1})_2(\mathbf{o-3.2})_2$.⁹⁹ This was ascribed to the high propensity of the phenothiazine-based ligand to form interlocked double cages.¹³³ This study further revealed that this approach of sterically modifying the donor groups is not independent from the backbone geometries, in particular when the latter differs in the structure (with acridone being flat and phenothiazine being significantly bent along the tricyclic aromatic system) in that the degree of $\text{Pd}(\text{donor})_4$ helicity, as dictated by the methylated pyridines, will be mechanically linked to the backbone-based central geometry of such a cage.

Crowley and coworkers exploited attractive interactions around the coordination sites to steer integrative self-sorting. They designed ligands **3.4** and **3.5**, differing solely in the presence or absence of amino groups on the pyridines. The combination of **3.4** with homoleptic cage $\text{Pd}_{23,54}$ led to the



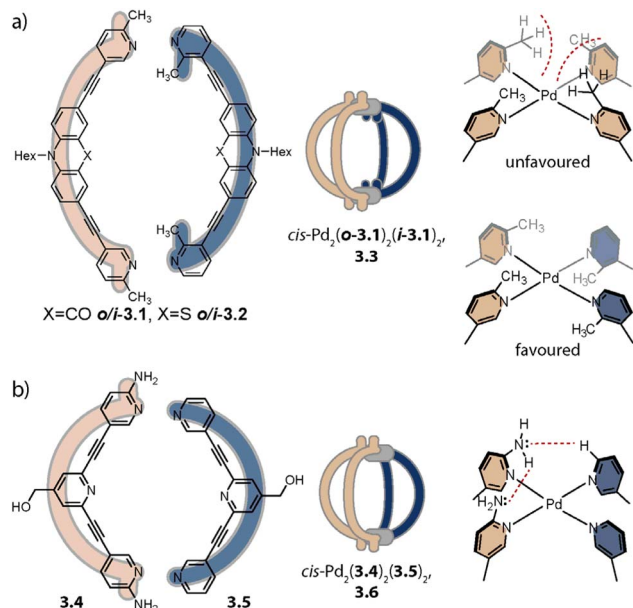


Fig. 11 Integrative self-sorting to *cis*-Pd₂A₂B₂-type cages using the CSE approach. (a) Repulsive steric interactions between the picolyl methyl groups destabilize a homoleptic with respect to a heteroleptic coordination environment resulting in **3.3** as the thermodynamic product. (b) Attractive interactions between the donor moieties kinetically stabilize cage **3.6**.

formation of cage *cis*-Pd₂(**3.4**)₂(**3.5**)₂, **3.6** for which the stereo configuration was assigned based on DFT computations (Fig. 11b). Preference for the *cis*-configuration was rationalized with the possibility for H-bond interactions between adjacent amino groups and additional H-bonds to the acidic α -protons of **3.5**. In line with DFT computations that predict homoleptic Pd₂(**3.4**)₄ to be the most thermodynamically stable system, heteroleptic *cis*-Pd₂(**3.4**)₂(**3.5**)₂ could not be accessed when the ligands, followed by the Pd(cations), or the preformed homoleptic cages were combined, *i.e.* it is a metastable, kinetic product.¹⁵⁴

Aside from steric constraints, geometrical principles have also been exploited for the formation of heteroleptic assemblies. Examples of this include metallacycles with *cis*-protected Pd(II) or Pt(II) by Stang¹⁵⁸ and Cu(II)-based coordination cages by Zhou,¹⁵⁹ both exploiting different ligand bent angles, and Pd₁₂A₁₂B₁₂ spheres by Fujita whereby A and B possess different lengths.¹⁶⁰

Concerning Pd_{*n*}L_{*2n*} cages, we introduced in 2016 a strategy exploiting complementary ligand geometries for achieving control over self-sorting and coined it “shape complementary assembly” (SCA).^{80,103} Ligands with convergent (**1.16b** or **2.55**) and divergent (**2.11**) binding vectors are combined to yield a single heteroleptic assembly in a non-statistical fashion (Fig. 12). Matching binding angles and suitable lengths allow for a square-planar coordination environment of the Pd(II) nodes, *e.g.* in *cis*-Pd₂(**1.16b**)₂(**2.11**)₂, **3.7** and *cis*-Pd₂(**2.55**)₂(**2.11**)₂, **3.8** without significant conformational strain. The severe twisting of the homoleptic helicate Pd₂(**1.16b**)₄ in combination with the high nuclearity of the homoleptic ring Pd₄(**2.11**)₈ drives

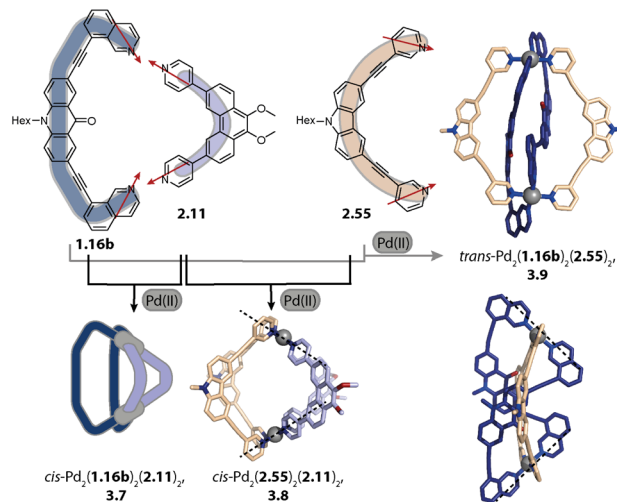


Fig. 12 Integrative self-sorting to Pd₂A₂B₂-type cages using the SCA approach. Shape complementarity between ligands **1.16b** or **2.55** and **2.11** permits selective formation of **3.7** and **3.8**, respectively. Combination of ligand **1.16b** with **2.55** leads to *trans*-configured **3.9**. The donor vectors are highlighted with red arrows and the tilted Pd(II) coordination planes with dashed lines.

the exergonic formation of **3.7**, both from an enthalpic and an entropic point of view. Addition of ligand **1.16b** to heteroleptic **3.8** results in transformation into **3.7** due to better shape complementarity of ligands **1.16b** and **2.11**. The complementarity of the binding angles leads to tilted Pd(donor)₄ planes and therefore to a bent cavity shape. This translates into a higher guest binding affinity of **3.7** for bent as compared to linear disulfonate guests (opposite to a straight, homoleptic cage of comparable size). Unexpectedly, the combination of **1.16b** and **2.55** also resulted in the formation of a defined species, *trans*-Pd₂(**1.16b**)₂(**2.55**)₂, **3.9**, in which ligands **1.16b** adopt an S-shaped conformation and cross through the centre of ring Pd₂(**2.55**)₂ (Fig. 12; compare also to its homoleptic analogue in Fig. 3d). The SCA approach proved to be widely applicable and was exploited for the synthesis of amphiphilic cages,¹⁶¹ CPL-active cages,^{162,163} tetranuclear cages,¹⁶⁴ and modular cage libraries for binding of phosphate esters.¹⁶⁵

Recently, Jelfs Lewis and coworkers reported a joint theoretical and experimental approach aiming at predicting the integrative self-sorting of convergent and divergent ligands to give *cis*-Pd₂A₂B₂ cages. From a theoretical perspective, the authors compared (a) the geometrical compatibility of ligand pairs and (b) the energy differences between homoleptic and heteroleptic species. The former yielded reasonable results in the sense that ligands for which integrative self-sorting was experimentally observed exhibited good geometrical compatibility. However, favourable geometric parameters did not ensure exclusive integrative self-sorting. Additionally, the relative energies of the overall assemblies were not congruent with the experimentally observed self-sorting outcome. These theoretical investigations were performed on single, static molecules, *i.e.* solvent and counter anion effects as well as entropic contributions were omitted in this simplified approach.¹⁶⁶ In



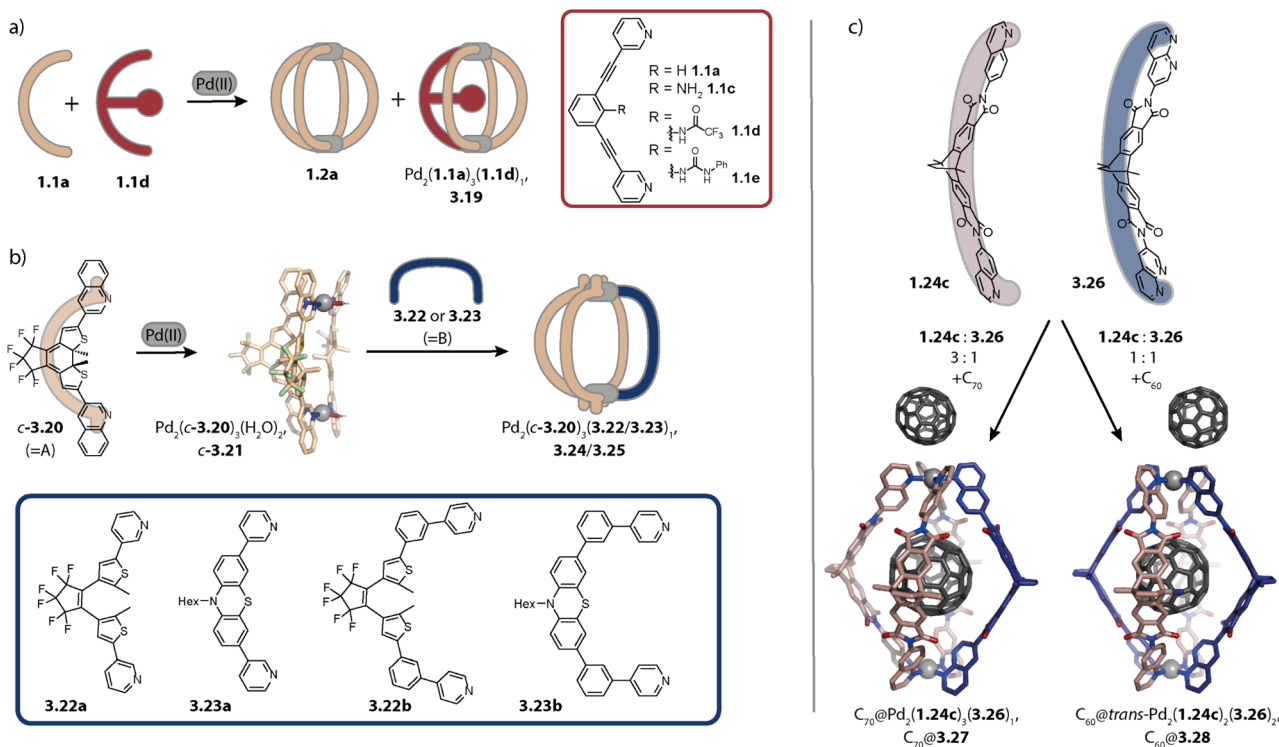


Fig. 14 Strategies for the self-assembly of Pd₂A₃B-type cages. (a) Installation of endohedral bulk with a particular size allows for assembly to **3.19** along with the homoleptic cage **1.2a**. (b) The CSE approach is applied for yielding bowl **c-3.21** with “free” coordination sites that can be bridged by ligands **3.22a–3.23b** lacking steric bulk at the donor moiety. (c) The CSE approach is expanded by attractive secondary interactions. Stoichiometry control then yields Pd₂A₃B₁ **3.27** and *trans*-Pd₂A₂B₂ cages **3.28** with encapsulated fullerenes.

sites. This strategy proved robust across different fourth ligands and hence allowed for the synthesis of a series of cages Pd₂(**c-3.20**)₃(**3.22a–3.23b**)₁, **3.24a–3.25b**. Depending on the linker (no linker in **3.22a** and **3.23a** vs. a phenyl linker in **3.22b** and **3.23b**), the cages possess different cavity sizes which translates into size selective guest encapsulation.¹⁶⁹

Moreover, we realized the installation of a fourth ligand on bowl-fullerene complex C₇₀@**1.25** in 2022. Again, coordination of a fourth ligand **1.24c** with isoquinoline donor moieties is unfavoured due to steric repulsion at the coordination site. In contrast, when a similar ligand with naphthyridine donors **3.26** is employed, C₇₀@Pd₂(**1.24c**)₃(**3.26**)₁, C₇₀@**3.27** is obtained (Fig. 14c, left). In addition to the reduced steric bulk of the naphthyridine in comparison to isoquinoline, the nitrogen lone pairs of the former allow for attractive interactions with the hydrogens of the adjacent donor groups of **1.24c**. Importantly, the CSE approach alone was not sufficient here for the clean formation of **3.27** in the absence of fullerene C₇₀ as a template. Additionally, ligands **1.24c** and **3.26** were combined in a 1 : 1 ratio with fullerene C₆₀ as a template, whereby C₆₀@*trans*-Pd₂(**1.24c**)₂(**3.26**)₂, C₆₀@**3.28** was obtained (Fig. 14c, right). The combination of the CSE approach and templation by fullerene also allowed for the assembly of heteroleptic bowls: when acridine-equipped ligand **1.24d** forming Pd₂(**1.24d**)₂(CH₃CN)₄ rings was combined with naphthyridine-equipped ligand **3.26** in the presence of C₆₀ or C₇₀, host-guest complexes C₆₀/C₇₀@*trans*-Pd₂(**1.24d**)₂(**3.26**)₁(CH₃CN)₂ were selectively

obtained. Here, naphthyridyl ligand **3.26** bridges one side of the ring Pd₂(**1.24d**)₂.¹⁵⁷

In 2024, Zhang and coworkers combined the CSE approach with the incorporation of endohedral steric bulk to form Pd₂A₃B₁ cages with varying cavity volumes. Picoly ligand **3.29** was combined with Pd(II) in a 3 : 2 ratio for obtaining metastable bowl Pd₂(**3.29**)₃(CH₃CN)₂, **3.30** (Fig. 15, top). In the next step, pyridine ligands **1.1a**, **1.1d**, **1.1f**, **1.1g**, and **1.1h** were added as ligand B to form cages Pd₂(**3.29**)₃B₁, **3.31**. While this stepwise synthesis allowed for obtaining the peculiar heteroleptic cages with sufficient selectivity, a one-pot reaction of the ligands with Pd(II) led to a mixture of different assemblies. For ligand B, possessing endohedral bulk (**1.1d**, **1.1f**, **1.1g**, and **1.1h**), the heteroleptic cage **3.31** was, however, the main species. This showcases how the interplay of two steric control elements can steer integrative self-sorting. Noteworthy, the accessible cavity volume decreases depending on ligand B in the order **1.1a** > **1.1f** > **1.1g**. This was shown to result in an alteration of the host-guest properties.¹⁷⁰ Wang and coworkers also combined coordination sphere engineering and endohedral steric bulk for realizing Pd₂A₂B₂ and Pd₂A₃B cages in a pathway-dependent approach.¹⁷¹

4.3. Pd₂A₂B₁C₁ cages

Zhang and coworkers reported in 2023 the first example of a heteroleptic Pd₂A₂B₁C₁-type cage incorporating three distinguishable ligands. Similar to their strategy for the synthesis of



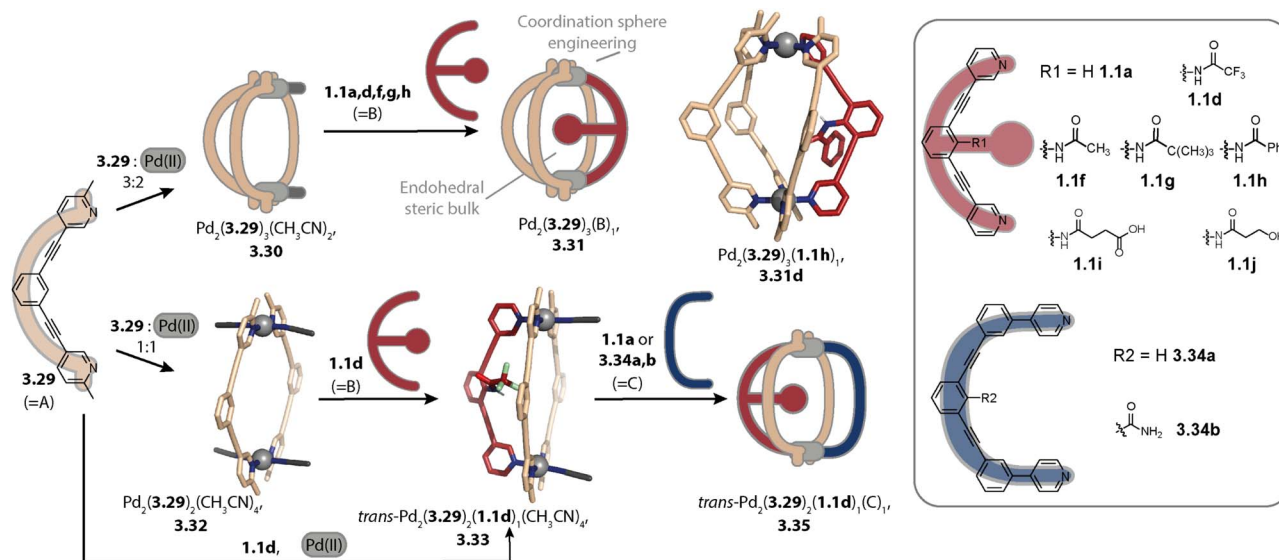


Fig. 15 Synthesis of Pd₂A₃B₁ (top) and Pd₂A₂B₁C₁ cages (bottom) through the incorporation of two steric control elements located at the donor moieties (CSE approach) and within the cage cavity. Picolyl ligand 3.29 assembles to form (metastable) bowl 3.30 or ring 3.32 depending on the ligand-to-metal stoichiometry. Bowl 3.30 can be converted into cage 3.31 with a fourth ligand lacking bulk at the donor site. The combination of ring 3.32 with a ligand equipped with endohedral bulk (shown for 1.1d) or the ligands with Pd(II) yields heteroleptic bowl 3.33 which can be converted into Pd₂A₂B₁C₁-type cage 3.35.

Pd₂A₃B₁-type cages, the CSE approach was merged with the installation of endohedral steric bulk. The combination of picolyl ligand 3.29 with Pd(II) in a 1:1 ratio led to ring Pd₂(3.29)₂(CH₃CN)₄, 3.32 (Fig. 15, bottom). Next, one side of the ring was bridged by ligand 1.1d, acting as a ligand “B”. While ligand 1.1d is of similar shape to 1.1a it possesses endohedral steric bulk, hampering coordination of a second equivalent. Thus, a heteroleptic bowl *trans*-Pd₂(3.29)₂(1.1d)₁(CH₃CN)₂, 3.33 is obtained. Noteworthy, bowl 3.33 can also be assembled directly from a mixture of ligands 3.29 and 1.1d with Pd(II), showing that its formation is pathway-independent. Lastly, a sterically less demanding ligand, such as unsubstituted 1.1a, can be implemented as ligand “C” to yield *trans*-Pd₂(3.29)₂(1.1d)₁(1.1a)₁, 3.35. The authors showed that different endohedral functionalities can be incorporated into the heteroleptic cage by changing the appended group in ligand B (1.1i: carboxylic acid and 1.1j: hydroxy group). Furthermore, the cavity can be enlarged by employing larger 3.34a as ligand C. On the other hand, a second endohedral moiety can be incorporated by functionalization of ligand C (3.34b).¹⁷²

In 2024, we combined the SCA approach with either inter-ligand interactions or with stoichiometry control to steer the selective formation of *cis*- or *trans*-Pd₂A₂B₁C₁ cages, respectively. In accordance with the SCA approach, ligand 1.24a assembles with ligands 2.33a or 2.33d to form *cis*-Pd₂(1.24a)₂(2.33)₂ cages. While ligands 2.33a and 2.33d feature similar geometries, their electronic properties diverge: fluorene-based ligand 2.33d possesses two CH₃ groups while fluorenone-based ligand 2.33a has a planar π-surface. CH₃⋯π interactions between central parts of neighbouring ligands 2.33a and 2.33d foster selective formation of *cis*-Pd₂(1.24a)₂(2.33a)₁(2.33d)₁, 3.36 when the ligands are combined in a ratio of 2 : 1 : 1 with Pd(II) (Fig. 16a).

For this, we coined the term “adjacent backbone interactions” (ABIs). Small ligands 2.33 also show shape complementarity to ligand 1.29 with strongly inward pointing donor vectors. Therefore, when 1.29, 2.33d, and 1.24a are combined with Pd(II) in a 1 : 1 : 2 : 2 ratio, *trans*-Pd₂(1.24a)₂(1.29)₁(2.33d)₁, 3.37 is obtained. That ligand 1.24a can adapt both, a conformation with rather convergent binding vectors, allowing for shape complementarity to ligands 2.33, and a conformation with nearly collinear binding vectors without considerable strain can be explained with its spring-like nature that enables adaptable lengths and binding angles.^{121,173}

4.4. Pd₂A₁B₁C₁D₁ cages

Utilizing the pathway-dependent self-assembly approach described earlier, Hiraoka and coworkers formed Pd₂A₃B₁, PdA₂B₁C₁, as well as the first Pd₂A₁B₁C₁D₁-type cage under kinetic control. Ring 3.11 was reacted with ligand 3.15 to yield intermediate bowl Pd₂(3.10)₂(3.15)₁(CH₃CN)₂, 3.38 that was then converted into Pd₂(3.10)₃(3.15)₁, 3.39 by the addition of a further equivalent of ligand 3.10 (Fig. 13). Reversing the last two steps did not lead to selective formation of 3.39 since homoleptic Pd₂(3.10)₄ can be formed. When intermediate bowl 3.38 was reacted with ligand 3.14, which is complementary in shape to ligand 3.15, cage *trans*-Pd₂(3.10)₂(3.14)₁(3.15)₁, 3.41 was obtained. 3.41 was also obtained when ligands 3.14 and 3.15 were added in the reverse order. Notably, this cage exhibited thermodynamic stability. This can be explained similar to the stability of cage 3.37: the ligands opposing each other either possess parallel binding vectors (ligand 3.10) or complementary geometries (3.14 and 3.15). In another vein, intermediate bowl 3.17 yielded cage *cis*-Pd₂(3.14)₂(3.15)₁(3.42)₁, 3.43 when ligand 3.42, only differing from ligand 3.15 by its



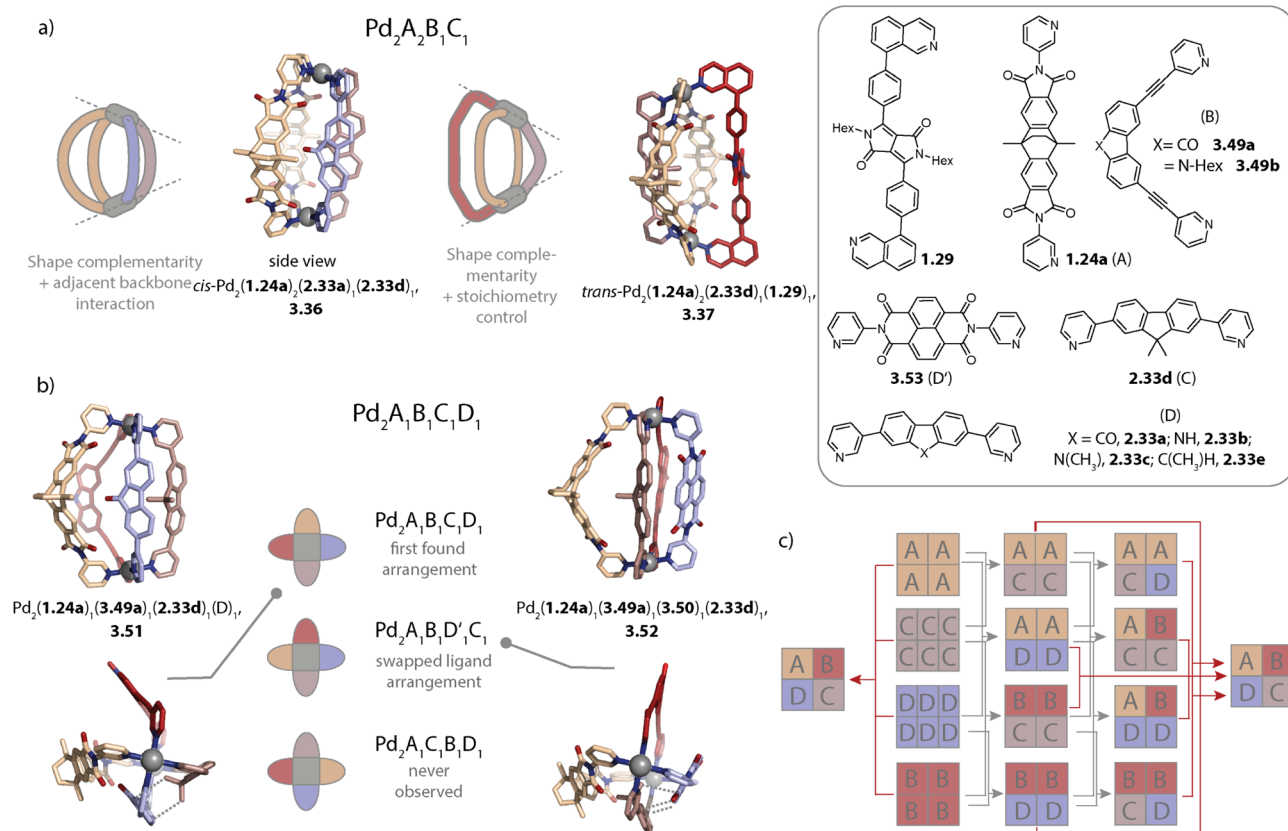


Fig. 16 Integrative self-sorting to Pd₂A₂B₁C₁-type and Pd₂A₁B₁C₁D₁-type cages under thermodynamic control. (a) Structures of *cis*-Pd₂A₂B₁C₁-type cage **3.36** and *trans*-Pd₂A₂B₁C₁-type cage **3.37** obtained by combining the SCA approach with the ABI approach or stoichiometry control, respectively. (b) Illustration of the three possible ligand arrangements in Pd₂A₁B₁C₁D₁-type cages and X-ray structures of realized pseudoisomers **3.51** and **3.52**, selectively formed as one of 55 homo- and heteroleptic dinuclear cages possible in principle. (c) All possible pathways for the synthesis of the Pd₂A₁B₁C₁D₁-type cage via cage-to-cage transformations.

exohedral substituents, was added. The culmination of the synthetic strategy represents the formation of a cage with four differentiable ligands: starting from ring **3.16**, bowl intermediate Pd₂(**3.14**)₁(**3.15**)₁(**1.1a**)₁(CH₃CN)₂, **3.44** was obtained upon addition of ligand **1.1a**. With ligand **3.10**, bowl **3.44** was then converted into Pd₂(**3.14**)₁(**3.15**)₁(**1.1a**)₁(**3.10**)₁, **3.45**. Except for cage **3.41**, heating of all assemblies discussed here resulted in a transformation into other, thermodynamically more stable compounds. In order to get insight into the factors influencing the respective cage stabilities, the authors determined the half-life of the cages and correlated them with different structural features. It was found that small deviations of the N⋯N distances of the ligands in the cages from their ideal distances are essential for high cage stability.¹⁶⁸

Finally, we published the first report on the thermodynamically controlled formation of Pd₂A₁B₁C₁D₁-type cages in 2024. Based on our knowledge on the shape complementarity of larger ligands **1.24a** (A) and **3.49** (B) with smaller ligands **2.33d** (C) and **2.33a** (D) as well as favourable inter-ligand interactions between the latter (ABI approach), we combined all four ligands with Pd(II) cations in a 1 : 1 : 1 : 1 : 2 ratio. Intriguingly, out of 55 possible dinuclear species, one single assembly incorporating all four ligands was exclusively obtained, both in solution and

the solid state. Considering the above-mentioned driving forces, two different isomers, namely Pd₂(**1.24a**)₁(**3.49a**)₁(**2.33d**)₁(**2.33a**)₁, **3.51** (Pd₂A₁B₁C₁D₁) and Pd₂(**1.24a**)₁(**3.49a**)₁(**2.33a**)₁(**2.33d**)₁ (Pd₂A₁B₁D₁C₁) would have been feasible whereby X-ray structure analysis revealed that only the former is formed (Fig. 16b). Concerning the observed stoichiometry, presumably, the incorporation of one equivalent of ligand **1.24a** is favoured due to its ability to adapt its N⋯N distance to match the other ligands. Fluorenone-based ligand **3.49a** could be replaced by carbazole-based ligand **3.49b**, possessing a similar shape and flexibility. Furthermore, we showed that ligands **2.33b**, **2.33c**, and **2.33e**, based on carbazole or fluorene, can likewise serve as ligand D since they also offer a π-surface for adjacent backbone interactions with ligand **2.33d**. In contrast, dimethyl fluorene ligand **2.33d** C proved to be vital for the selective outcome since substitution by ligands **2.33b**, **2.33c**, or **2.33e** which offer only up to one CH₃ group in the central backbone position resulted in the formation of mixtures of different species. Interestingly, when NDI-based ligand **3.53** was applied as ligand D, a swapped ligand arrangement was observed. In this case, ligand **2.33d** offers its π-surface for CH⋯π-interactions with ligand **3.53**. Probably, the alteration of the orientations of ligand C and D affects the shape



complementarity with ligands A and B in a way that now favours the formation of pseudo-isomeric cage 3.52. Noteworthy, cage $\text{Pd}_2\text{A}_1\text{B}_1\text{C}_1\text{D}_1$ could also be obtained in a stepwise fashion through several possible cage-to-cage transformation pathways, starting from the respective homoleptic cages, again supporting its characteristics as the final thermodynamic assembly product (Fig. 16c).¹²¹

4.5. $\text{Pd}_n\text{A}_x\text{B}_y$ assemblies with $n > 2$

Recently, we reported the first trinuclear heteroleptic assemblies $\text{Pd}_3\text{A}_3\text{B}_3$. By applying the SCA approach, the long, convergent ligand 4.1 was combined with the small, divergent ligand 3.46a. This resulted in the formation of three-fold symmetric *syn-cis*- $\text{Pd}_3(\text{3.46a})_3(\text{4.1})_3$, 4.3a assembly (Fig. 17). Here, *cis* refers to the ligand arrangement around the metal node and *syn* to the ligand orientation relative to the Pd_3 plane, *i.e.* the assembly is overall bowl-shaped. Modification of the central benzene moiety (ligands 4.2b–e) allowed for obtaining endohedrally functionalized rings 4.3b–e. Noteworthy, when uncharged and small functionalities were incorporated, the self-assembly to $\text{Pd}_3\text{A}_3\text{B}_3$ -type rings was not disturbed. On the other hand, when the sterically more demanding *tert*-butyl ester modified ligand 4.2f was employed, a mixture of trinuclear and tetranuclear assemblies was obtained. A further increase in steric bulk (phenyl ester ligand 4.2g) led to the tetranuclear ring as the sole species. In a similar fashion, functionalization with a nitro group led to higher nuclearity ring $\text{Pd}_4(\text{4.1})_4(\text{4.2h})_4$, 4.5h. Presumably, this can be traced back to reduced electrostatic repulsion between the nitro oxygens of ligands 4.2h when occupying neighboring edges in the larger assembly as compared to a tentative three-membered ring. Aside from the incorporation of steric bulk and charges, the formation of a $\text{Pd}_4\text{A}_4\text{B}_4$ -type assembly could also be steered by modification

of the bent angle of the small ligand. When the central benzene moiety in 3.46a ($\alpha = 120^\circ$) was substituted by a thiophene or a selenophene moiety, the bent angle was increased to 149° (4.4a) and 151° (4.4b), respectively. In accordance with the trends observed for homoleptic assemblies (Fig. 4), this led to an increase in nuclearity, here to a $\text{Pd}_4\text{A}_4\text{B}_4$ -type ring. Interestingly, NMR analyses in combination with DFT computations suggest that for most of the four-membered rings investigated, more precisely the rings formed with 4.4a, 4.4b, and 4.2g, a saddle-shaped conformation is favoured over a bowl-shaped one.⁴¹

In 2019, we reported the self-assembly of pill-shaped tetranuclear species by exploiting the charge separation strategy, widely utilized by Stang and coworkers,¹⁵ combined with our CSE approach. As discussed in the preceding section, the “free” coordination sites of quinoline-based Pd_2A_3 -type bowls (such as 1.25, Fig. 18a) can be occupied by ligands lacking steric bulk close to the donor site. When instead of a convergent bis-pyridyl ligand a linear terephthalate 4.6 is employed, two bowls are bridged, giving rise to pill-shaped $\text{Pd}_4(\text{1.7})_6(\text{4.6})_2$, 4.7. The negative charges of the terephthalates help compensate for the repulsion stemming from the close proximity of multiple Pd(II) cations. Additionally, the dimer-fullerene complexes $2\text{C}_{60}@4.7$ or $2\text{C}_{70}@4.7$ could also be accessed upon dimerization of pre-formed bowl-fullerene complexes $\text{C}_{60}@1.25$ and $\text{C}_{70}@1.25$.⁷⁶

In another vein, Hiraoka and coworkers recently reported pill-shaped $\text{Pd}_4\text{A}_4\text{B}_2\text{C}_2$ -type assemblies *via* dimerization of two $\text{Pd}_2\text{A}_2\text{B}$ heteroleptic bowls under kinetic control. Starting from bowl 3.38, whose synthesis is described in Chapter 4.4, the authors added long or short bis-pyridyl ligands 3.46 with diverging binding angles. Similar to our approach, pill-shaped dimer $\text{Pd}_4(\text{3.10})_4(\text{3.15})_2(\text{3.46})_2$, 3.47 was obtained, albeit here with three different bis-pyridyl ligands and higher positive charge (Fig. 13).¹⁶⁸

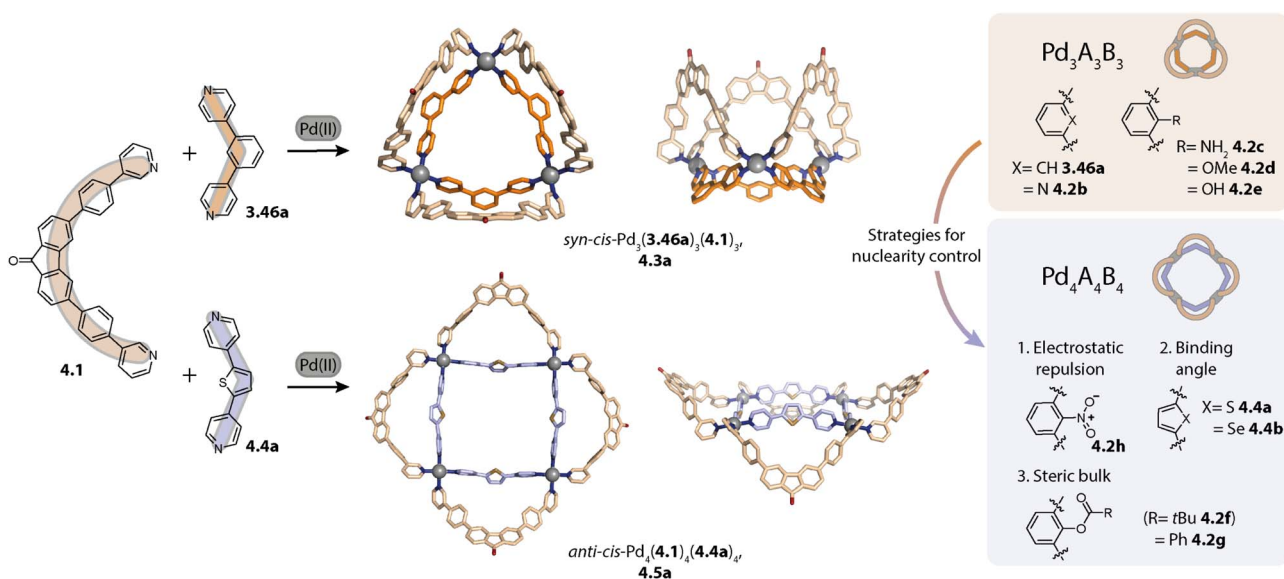


Fig. 17 Self-assembly of $\text{Pd}_3\text{A}_3\text{B}_3$ -type and $\text{Pd}_4\text{A}_4\text{B}_4$ -type heteroleptic rings. (Left) Ligand 4.1 self-assembles with ligand 3.46a to form bowl-shaped 4.3a (top) and with ligand 4.4a to form saddle-shaped 4.5a (bottom). (Right) The ring nuclearity can be increased upon modification of the smaller ligand either through (1) the incorporation of a nitro substituent, (2) an increase in the bent angle, or (3) installation of steric bulk.



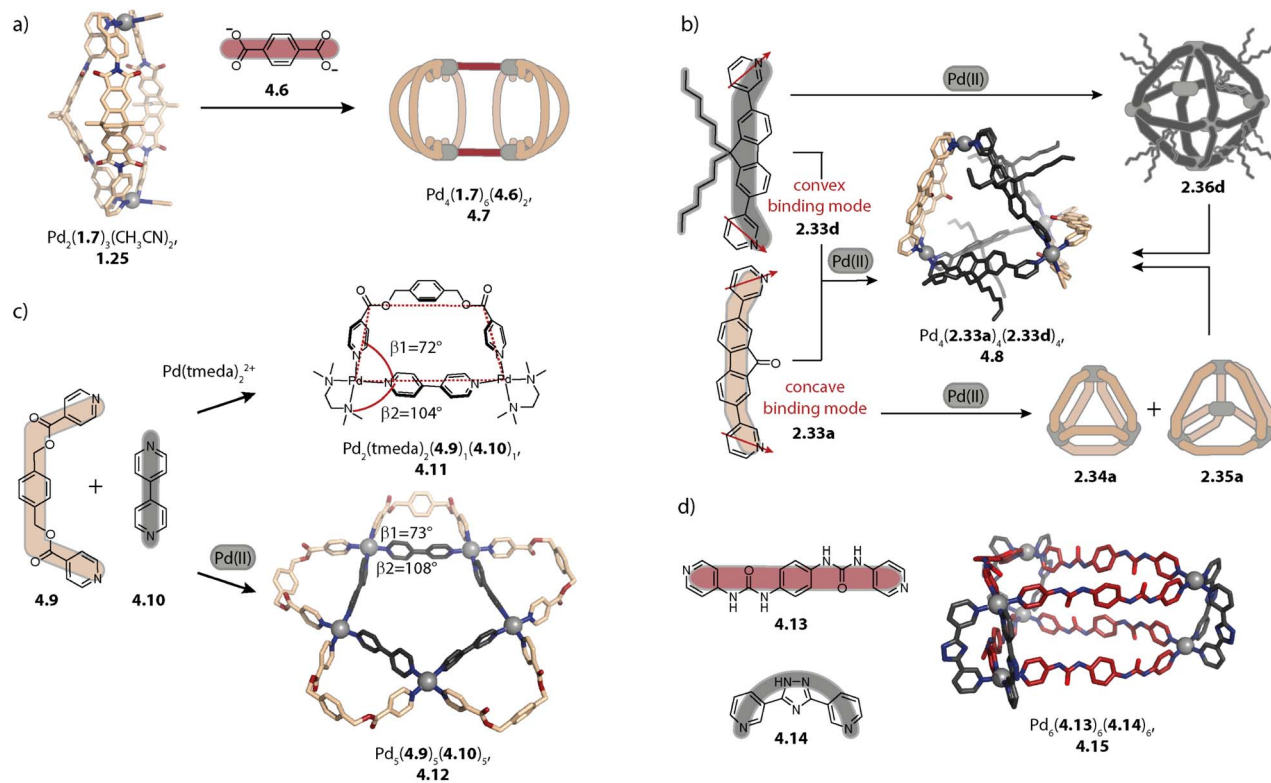


Fig. 18 (a) Pill-shaped dimer **4.7** is obtained by bridging two counts of bowl **1.25** with dicarboxylate ligand **4.6**. (b) Fluorene-derived ligands **2.33a/d** self-assemble either to form a mixture of ring **2.34** and tetrahedron **2.35** or to form an octahedron **2.36d** depending on the backbone-centred steric bulk. Combination of **2.33a** and **2.33d** with Pd(II) results in integrative self-sorting to heteroleptic tetrahedron **4.8**. (c) Self-assembly of flexible ligand **4.9** with linear ligand **4.10** gives rise to trapezium-shaped complex **4.11** when a *cis*-protected Pd(II) source is employed and to five-membered ring **4.12** with unprotected Pd(II) (β_1 and β_2 are exterior and interior angles of the truncated pentagram). (d) Trigonal prismatic assembly **4.15** formed with urea-functionalized ligand **4.13** and ligand **4.14**.

In 2021, we reported a strategy for yielding heteroleptic tetrahedra by exploiting the role of backbone-centred steric bulk (as a special variant of the ABI approach). As described in Chapter 3.2, fluorene-derived ligands **2.33** can assemble to form rings, pseudo-tetrahedra, and octahedra, depending on the steric bulk installed at the backbone. While ligand **2.33a** with low steric demand assembles to form a mixture of ring **2.34a** and tetrahedron **2.35a**, dihexyl-decorated ligand **2.33d** forms entropically disfavoured homoleptic octahedron **2.36d** (Fig. 18b). In the latter, the distance between the ligands is increased and they adopt a concave binding mode while the steric bulk is oriented to the exterior. Upon careful inspection of the ligand conformations and arrangement in tetrahedron **2.35a** (Fig. 6), we suspected that the assembly could serve as a platform for combining sterically demanding ligands, sitting on the singly bridged edges, with ligands of low steric demand occupying the doubly bridged edges. Indeed, a combination of either ligands **2.33a** and **2.33d** with Pd(II) or combination of the preformed homoleptic assemblies **2.34a** and **2.35a** with **2.36d** resulted in the selective formation of tetrahedron Pd₄(**2.33a**)₄(**2.33d**)₄, **4.8** with the expected ligand arrangement (Fig. 18b).¹¹⁹

Yoshizawa, Chand, and coworkers reported the first and, until now, sole pentanuclear heteroleptic assembly in 2017 (elected as the molecule of the year in ACS Chemical & Engineering News). In the first step, the authors combined the rather flexible ligand **4.9** with linear ligand **4.10** and a *cis*-protected Pd(II) source, (tmeda)Pd(II). A trapezium-shaped complex Pd₂(tmeda)₂(**4.9**)₁(**4.10**)₁, **4.11** with an exterior angle β_1 of 72° and an interior angle β_2 of 104° was obtained, serving as a promising precursor to a higher order assembly (Fig. 18c).¹⁷⁴ It was then anticipated that the combination of the same two ligands with unprotected Pd(II) should result in a circular arrangement of fused trapezium-shaped complexes **4.11**. Comparing the angles observed in **4.11** with the angles in tentative triangular, square, and pentagonal arrangements of annelated trapezia suggested that, from a purely geometrical point of view, the pentagonal species should be favoured. Indeed, the combination of **4.9** and **4.10** with [Pd(CH₃CN)₄](BF₄)₂ in DMSO resulted in pentanuclear species Pd₅(**4.9**)₅(**4.10**)₅, **4.12** which can be described as a truncated pentagram. Here, the interior angle reaches 108°, similar to the one of a regular pentagon. Noteworthy, ring **4.12** could also be accessed *via* different pathways, more precisely upon combination of the two homoleptic species or upon addition of one of



the ligands to the homoleptic assembly of the respective other ligand, supporting that **4.12** is the thermodynamic product.¹⁷⁵

In 2016, Mukherjee and coworkers reported a trigonal prismatic Pd₆A₆B₆-type cage. First, the authors showed that urea-functionalized bis-pyridyl ligand **4.13** assembles with *cis*-protected (tmeda)Pd(II) to form a molecular triangle Pd₃(tmeda)₃(**4.13**)₃. Aiming at a three-dimensional assembly, ligand **4.13** was combined with unprotected Pd(II) and short “clipping” ligand **4.14**. In the resulting assembly Pd₆(**4.13**)₆(**4.14**)₆, **4.15**, in total six clipping ligands **4.14** bridge two triangles Pd₃(**4.13**)₃ (Fig. 18d). Urea-functionalized compounds are generally prone to self-association *via* hydrogen bonds which limits their capability of substrate recognition and hence their catalytic applicability. The large inter-ligand distances in the triangular prism prohibit such self-quenching interactions, rendering the assembly suitable for the recognition of nitro-olefins. Therefore, triangular prism **4.15** could be used as a heterogeneous catalyst for promoting Michael-additions and Diels–Alder reactions in water.¹⁷⁶

In 2021, Severin and coworkers achieved selective assembly of Pd₄A₄B₄-type tetrahedral, Pd₆A₆B₆-type trigonal prismatic, and Pd₈A₈B₈-type tetragonal prismatic architectures by

variation of the binding angle of one of the employed ligands. Heteroleptic assembly formation was investigated with ligand **4.16**, possessing collinear binding vectors. The ligand is pre-disposed for the formation of homoleptic assemblies as its homoleptic cage Pd₂(**4.16**)₄, **4.17** is strained due its small Pd···Pd distance, amongst other factors (Fig. 19a). The combination of **4.16** with ligand **3.46a**, possessing a bent angle of 120° led to selective formation of pseudo-tetrahedral Pd₄(**4.16**)₄(**3.46a**)₄, **4.18a**. In accordance with the binding modes observed in homoleptic tetrahedra, more open ligand **3.46a** occupies the singly bridged edges while two ligands **4.16** form macrocyclic motif Pd₂(**4.16**)₂. Interestingly, a similar outcome was achieved when alkyne-spaced ligand **3.46b** was employed; however, ligand **4.16** could not be replaced by its alkyne spaced analogue **1.1a**. This apparent disparity can be explained with the higher strain of homoleptic **4.17** in comparison to the longer analogue **1.1a**. The combination of wider ligand **4.4a**, based on thiophene, with ligand **4.16** resulted in the formation of a trigonal prismatic cage Pd₆(**4.4a**)₆(**4.16**)₆, **4.19**. Similar to the assembly **4.15** reported by Mukherjee and coworkers, two counts of ligand **4.16** form macrocyclic moieties Pd₂(**4.16**)₂ that are bridged by ligands **4.4a**. Next, the authors further increased the binding

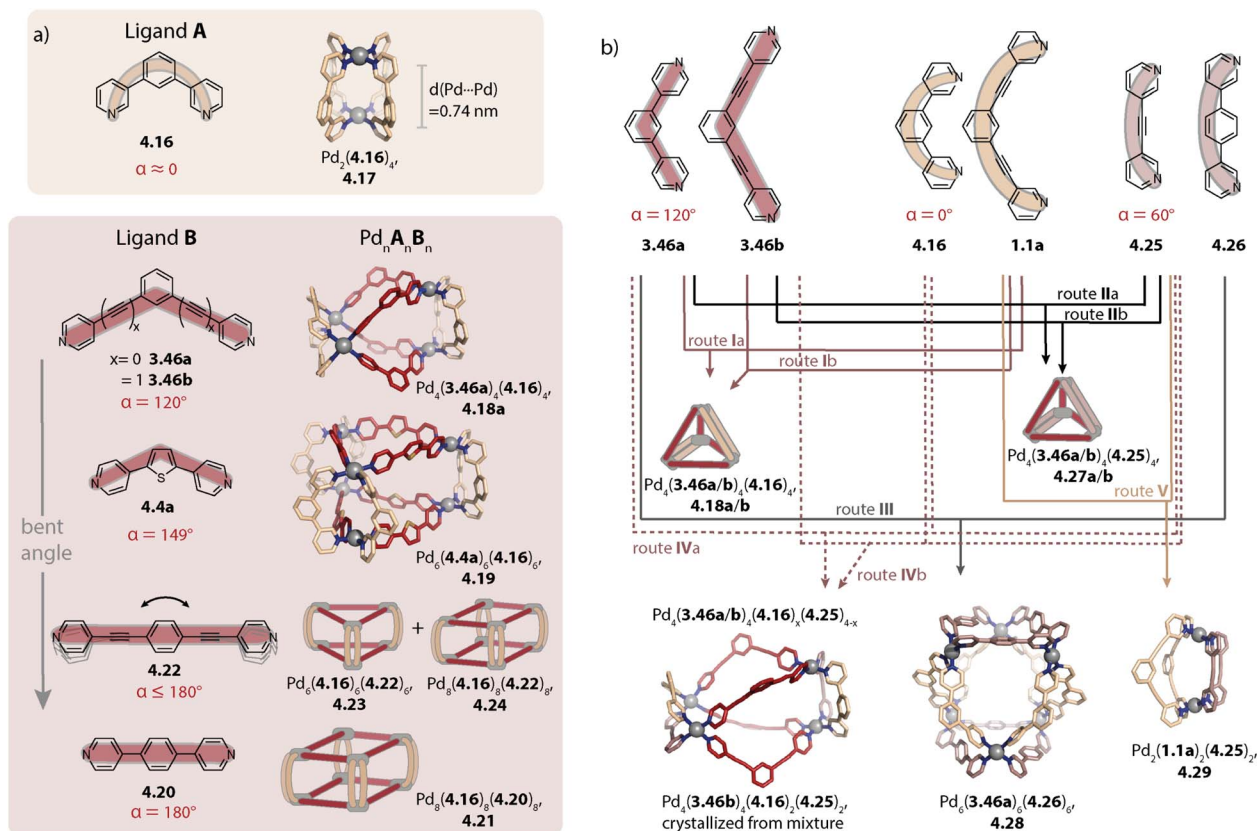


Fig. 19 (a) Integrative self-sorting to heteroleptic Pd_nA_nB_n-type assemblies ($n = 4, 6, 8$) through the combination of small ligand **4.16** with collinear binding vectors with ligands **3.46**, **4.4a**, **4.22**, and **4.20** possessing divergent vectors. The bent angle and the flexibility of the latter control the nuclearity. (b) Overview of heteroleptic assembly formation with binary or ternary mixtures of ligands with collinear vectors (**1.1a** and **4.16**), 60° binding angles (**4.26** and **4.25**), or 120° angles (**3.46**). Tetrahedral Pd₄A₄B₄ can be obtained from 0°-bent angle ligand **4.16** with 120°-bent angle ligands of different lengths (route I) or upon combination of one of the latter two with 60°-bent angle ligand **4.25** (route II). Tetrahedral Pd₄A₄B₄C_{4-x} can be obtained when ligands **3.46a** or **3.46b** are combined with ligands **4.16** and **4.25**. Trigonal antiprismatic assembly **4.28** is obtained with ligands **4.16** and **4.25**. Ligands **1.1a** and **4.25** self-assemble to form dinuclear cage **4.29**.



angle of the second ligand by applying linear ligands in which the *para*-pyridines were either bridged by a phenyl (**4.20**) or by a *para*-diethynylbenzene linker (**4.22**). From a purely geometrical point of view, the combination of a linear ligand with ligand **4.16** would lead to the formation of a tetragonal prismatic (cuboid) assembly. Indeed, the exclusive formation of $\text{Pd}_8(\mathbf{4.16})_8(\mathbf{4.20})_8$, **4.21** was observed for short, and thus rigid, ligand **4.20**. Conversely, the alkyne linkers endow ligand **4.22** with enough flexibility to allow for the formation of entropically favoured trigonal prism $\text{Pd}_6(\mathbf{4.16})_6(\mathbf{4.22})_6$, **4.23** along with tetragonal prism $\text{Pd}_8(\mathbf{4.16})_8(\mathbf{4.22})_8$, **4.24**.¹⁵⁶

Tetrahedral assemblies can be obtained upon the combination of 120° bent angle ligands **3.46** with ligand **4.16** (as just discussed, Fig. 19b, route I). Ligand **4.16** occupying the doubly-bridged edges can be replaced by ligand **4.25**, possessing an angle of 60° (route II). Excitingly, the combination of the two phenyl-spaced ligands **3.46a** and **4.26** led to novel species $\text{Pd}_6(\mathbf{3.46a})_6(\mathbf{4.26})_6$, **4.28**, a trigonal antiprismatic assembly (route III). Severin and coworkers accomplished this discovery by screening of a virtual combinatorial library. More precisely, the authors combined a variety of well-known ligands (**1.1a**, **1.8**, **2.24**, **2.55**, bis-pyridyl analogue of **3.1**, **3.46a**) with substoichiometric amounts of a Pd(II) salt in order to favour formation of particularly stable assemblies. Examination of the X-ray crystal structure reveals that the peculiar ligand arrangement in **4.28** allows for a nearly ideal square planar coordination geometry of the Pd(II) nodes. For examining the generality of the assembly of hexanuclear heteroleptic cages, the authors employed ligands with similar angles albeit with different lengths and bulk. Hence, instead of ligand **4.26**, clathrochelate ligand **2.46d** was employed and ligand **3.46a** was replaced by an analogue with similar clathrochelate moieties between the 1,3-benzene-core and the *para*-pyridines. Interestingly, this led to the formation of a hexanuclear prismatic architecture instead of an antiprismatic one as observed in **4.28**.¹⁷⁷ In another vein, the authors showed that ligand **1.1a** assembles with ligand **4.25** to form dinuclear cage *cis*- $\text{Pd}_2(\mathbf{1.1a})_2(\mathbf{4.25})_2$, **4.29**. The tilted PdN₄ planes point towards shape complementarity between the two ligands. This is achieved by the bending of the ligands owing to their flexible alkyne spacers. In order to further increase the complexity of the system, the authors aimed at incorporating all three kinds of ligands (*i.e.* with 0°, 60°, and 120° bent angles) within one assembly. For this, the self-assembly was examined with the various ligand combinations (eight combinations and different ratios). Solely the combination of ligands **4.16** and **4.25** with either **3.46a** or **3.46b** afforded assemblies comprising all three ligands (routes IVa and IVb). More specifically, tetrahedra $\text{Pd}_4(\mathbf{3.46a/b})_4(\mathbf{4.16})_x(\mathbf{4.25})_{4-x}$ were obtained in which 120° ligand **3.46** occupies the singly bridged edges while ligands **4.16** and **4.25** are statistically distributed over the doubly bridged edges.²⁰⁸

Driven by the aim of preparing cages that carry multiple functionalities, vast progress has been achieved in the development of strategies for designing Pd(II)-based heteroleptic cages in the last decade. For favouring a heteroleptic coordination environment in dinuclear cages, the donor sites can be equipped with steric bulk or functionalities that allow for

attractive secondary interactions (coordination sphere engineering, CSE). Alternatively, repulsive or attractive interactions are installed in more central ligand positions, as in the adjacent backbone interaction (ABI) approach. In other examples, cavity-filling steric bulk is employed to steer the self-sorting outcome. The necessity for such ancillary groups can be omitted when suitably shaped ligands are used in the shape complementary assembly (SCA) strategy. The combination of different strategies was shown to allow for accessing dinuclear cages with up to four differentiable ligands. Similar strategies have been employed for designing heteroleptic cages of higher nuclearity. In this case, exohedral steric bulk was also added to control selective self-sorting. For $n = 6$ and $n = 8$, some topologies were accessed that could never be observed when only one kind of ligand was employed, namely trigonal prismatic and antiprismatic as well as tetragonal prismatic assemblies. In contrast to dinuclear cages, defined heteroleptic assemblies of higher nuclearity which carry more than two kinds of ligands have so far rarely been reported and thus pose a challenge for future investigations.

5. Orientational self-sorting

An alternative approach for accessing Pd_nL_{2n} assemblies of higher structural complexity is the use of ligands lacking the typical two-fold symmetry seen in all ligands discussed so far. On the one hand, this method can give rise to assemblies with Pd(II) ions sitting in differentiable coordination environments. On the other hand, it permits for the partitioning of functional groups within the cavity, holding promise for cages capable of binding low-symmetry guests. Most of the work reported so far focuses on the smallest family member, the Pd_2L_4 cage, where four isomers can be obtained depending on the relative arrangement (*H*: head and *T*: tail) amongst the ligands (Fig. 20, middle). Significant progress regarding the orientational self-sorting of asymmetric ligands was achieved in recent years by the groups of Lewis,^{179,187,189,193,197} Chand,^{191,192,194} and others^{180,181,186,188,196} from the experimental side and the group of Jelfs^{179,187,189} from the theoretical side.

In 2020, Chand and coworkers showed that ligand **5.1**, possessing an asymmetrical linker between the two pyridines, assembles to form a mixture of different isomers $\text{Pd}_2(\mathbf{5.1})_4$, **5.2** (Fig. 20, middle).¹⁷⁸ Later, the ligand design was picked up by Lewis, Jelfs, and coworkers who substituted one of the donor moieties with sterically more demanding picolyl (**5.1P**) or quinolinyl (**5.1Q**) moieties. Interestingly, the picolyl-based ligand **5.1P** shows a preference for the *HTHT-5.2P* isomer, while self-assembly of quinoline-based ligand **5.1Q** leads to a higher fraction of *HHHT-5.2Q* isomer when DMSO is used as the solvent. Addition of acetonitrile led to an increase in the fraction of *HTHT-5.2* for both ligands. Supported by DFT-computations, the authors suggested that the *HTHT-5.2* isomer is favoured from a purely structural point of view while the *HHHT*-isomer offers a suitable pocket for binding of H-bond acceptor solvents such as DMSO. The latter effect is of particular importance for the ligand with quinoline donors due to the polarized quinoline proton. Pleasingly, this behaviour was also



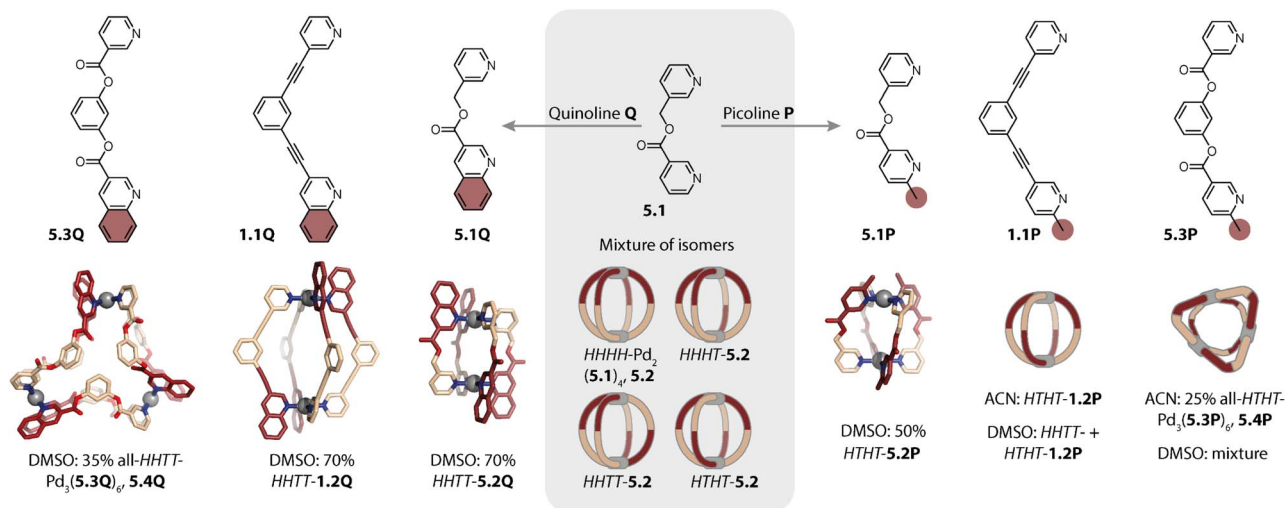


Fig. 20 Biasing the self-assembly of asymmetric ligands by introducing steric constraints on the donor site. (Left) For quinoline-equipped ligands the *HHTT*-isomer is favoured in DMSO. (Middle) Bis-pyridyl ligand **5.1** assembles to form a mixture of different isomers. (Right) For picolyl-equipped ligands, there is a bias towards *HTHT*-isomers.

observed for the self-sorting of ligands **1.1Q** and **1.1P**: the former shows a preference for the *HHTT*-**1.2Q** isomer while the latter assembles to form *HTHT*-**1.2P** in acetonitrile and to form a mixture of the two isomers in DMSO. Ligands **5.3Q** and **5.3P** assemble to form three-membered rings $\text{Pd}_3(\mathbf{5.3})_6$, **5.4**. Similar effects lead to a preference for all-*HHTT*- $\text{Pd}_3(\mathbf{5.3Q})_6$, **5.4Q** in the case of **5.3Q** in DMSO and for all-*HTHT*- $\text{Pd}_3(\mathbf{5.3P})_6$, **5.4P** in the case of **5.3P** in ACN.¹⁷⁹

In 2023, Lusby, Crowley, and coworkers investigated the effect of attractive interactions within the coordination sphere

on the orientational self-sorting. A selection of ligands, termed here **R1-1.1-R2** (Fig. 21a), with variable hydrophilicity has been synthesized: ligands **H-1.1-Alk3**, **H-1.1-Alk6**, and **H-1.1-Alk12** are equipped with a hydrophobic alkoxy chain on one of the pyridines and ligands **EG-1.1-H**, **EG2-1.1-H**, and **EG4-1.1-H** possess hydrophilic ethylene glycol chains of increasing length. Furthermore, ligands **EG-1.1-Alk3**, **EG2-1.1-Alk6**, and **EG4-1.1-Alk12** have been synthesized which are functionalized with alkoxy chains on one ligand end and ethylene glycol chains on the other end that are of varying length. Intuitively, DFT computations supported that for the **EG-1.1-Alk** ligands, the *HHHH*- $\text{Pd}_2(\mathbf{1.1})_4$ isomer, in which hydrophobic and hydrophilic chains are segregated, should be favoured. A similar isomer was favoured for **H-1.1-Alk** ligands, while the *HTHT*-isomer was predicted to be the most stable for the **EG-1.1-H**-type ligands. The predicted energy differences between the isomers were, however, small. In accordance with this, the authors observed the formation of mixtures of isomers of $\text{Pd}_2(\mathbf{1.1})_4$ when the ligands were combined with $\text{Pd}(\text{II})$. Hence, dispersion interactions were not sufficient here for steering the self-sorting towards a single isomer. The authors also equipped one of the ligand ends with 2-amino pyridine. As discussed in Section 4.1, this strategy allowed for the formation of (kinetically stable) *cis*- $\text{Pd}_2\text{A}_2\text{B}_2$ with the corresponding symmetrical ligands A and B. In a similar fashion, ligand **5.5a** assembled to form *HHTT*- $\text{Pd}_2(\mathbf{5.5a})_4$, **5.6a** (Fig. 21c). DFT computations supported that the orientational self-sorting was steered by H-bonding interactions between neighbouring amino groups and that the isomer was, similar to heteroleptic **3.6**, a kinetic product. When the amino groups were positioned farther away from the nitrogen donor, more precisely in the *meta*-position for ligand **5.5b**, a mixture of isomers was obtained, highlighting the importance of the intramolecular H-bonds for the selectivity.¹⁸⁰

H-bonding interactions for driving orientational self-sorting were also exploited by Natarajan and coworkers. In 2019, the

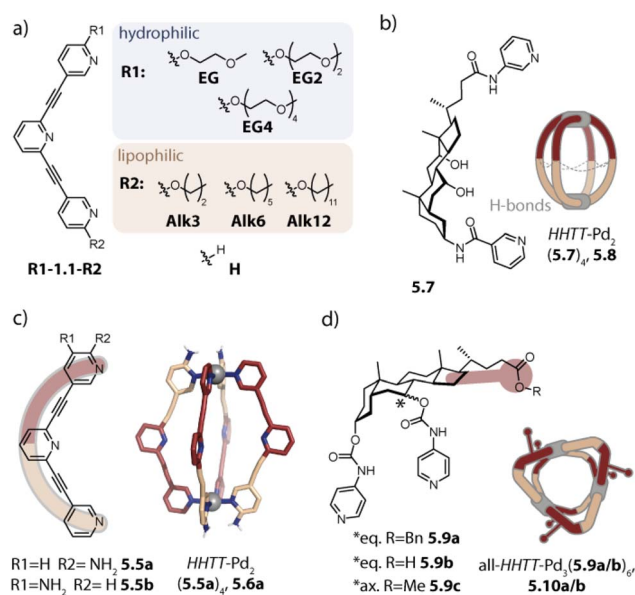


Fig. 21 (a) Design of ligands for an attempt to steer orientational self-sorting through dispersive interactions. (b) Self-sorting of ligand **5.7** to *HHTT*-**5.8** through backbone-centred H-bonds. (c) Steering self-sorting through H-bonds between amino substituents on the donor sites. (d) Ligands **5.9** and selectively formed all-*HHTT* isomer **5.10**.



authors reported dinuclear cages $\text{Pd}_2(5.7)_4$, **5.8** formed by ligand **5.7** which is based on cholic acid and possesses two hydroxyl groups on its flanks (Fig. 21b). The latter allows for inter-ligand hydrogen bonds in the helically twisted *HHHH*-**5.8** isomer. These interactions were vital for selectivity since removal of one or both hydroxy groups resulted in the formation of a mixture of isomers.¹⁸¹ Further recent reports by McTernan as well as Palma and coworkers on the stereoselective cage assembly from peptide-based bis-pyridyl ligands are noteworthy.^{182,183}

Rissanen and coworkers reported in 2015 on the self-assembly of bile acid-based ligand **5.9a**, having a multitude of chiral centres, to a single isomer of three-membered ring $\text{Pd}_3(5.9a)_6$, **5.10a** (Fig. 21d). NMR experiments in combination with DFT computations and chloride-induced disassembly studies point towards an all-*HHTT* arrangement of the ligands in **5.10a**.¹⁸⁴

In a later study, the authors employed ligand **5.9b**, which possesses, in contrast to ligand **5.9a** a free carboxyl group. Assembly $\text{Pd}_3(5.9b)_6$, **5.10b** could only be obtained through stepwise synthesis by first adding a Fe(III) or Cu(II) salt, followed by a transmetalation with Pd(II). Supported by ion mobility mass spectrometric experiments, the authors suppose that the ligands are arranged in an all-*HHTT* fashion, similarly to **5.10a**. The unprotected carboxyl group renders **5.10b** water soluble and amphiphilic. The authors also observed higher order self-assembly: addition of water to the cage (synthesized by transmetalation of the Cu(II) assembly) in DMSO led to aggregation and gel formation. Furthermore, investigation with transmission electron microscopy revealed evaporation-induced formation of hexagonal particles.¹⁸⁵

Very recently, some of the authors incorporated a bile acid which is an epimer to **5.9a/b**, into bis-pyridyl ligand **5.9c**. Importantly, since the pyridine moiety is attached *via* the epimeric hydroxy group, ligand **5.9c** has an altered binding angle. Assembly of **5.9c** with Pd(II) led to a mixture of $\text{Pd}_3(5.9c)_6$, $\text{Pd}_4(5.9c)_8$, $\text{Pd}_5(5.9c)_{10}$, and $\text{Pd}_6(5.9c)_{12}$ species.¹⁸⁶

Lewis and coworkers showed in 2020 that the asymmetry of ligand **5.11**, in which the two ligand arms differ only in one alkyne spacer, is not sufficient for steering orientational self-sorting to one defined species (Fig. 22a). More precisely, the *HHTT*- and the *HHHT*-isomer were predicted to be relatively similar in energy. The bias for the *HHTT*-isomer could be sufficiently increased by endowing one end of the ligand with a picolyl donor moiety (**5.12**). Hence, orientational self-sorting was achieved here through a combination of steric and geometrical constraints (principle I). In another fashion, selectivity for the *HHTT*-isomer was achieved by increasing the geometrical constraints: in **5.13** and **5.14**, the planes orthogonal to the donor vectors are parallel but have, however, a larger offset as compared to ligand **5.11** (principle II). In both cases, the *HHTT*-isomer was obtained within two hours at room temperature. Ligand **5.15** differs from ligand **5.13** in the alkyne spacer only, which increases the similarity between the two arms. Selective self-assembly to the *HHTT*-isomer was observed; however, the equilibration time was significantly longer as compared to the one for the self-assembly of ligands **5.13** and **5.14**.¹⁸⁷

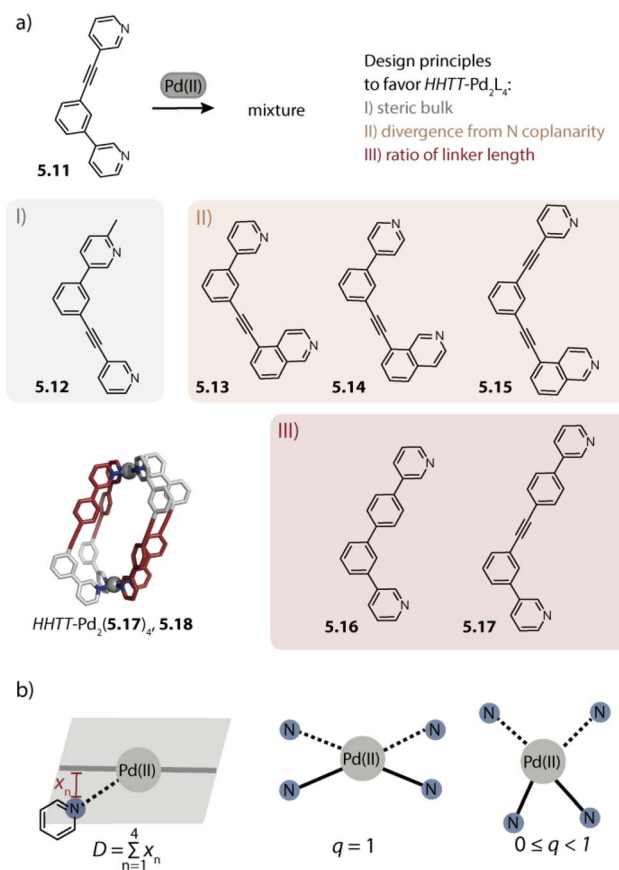


Fig. 22 (a) Modification of ligand **5.11** by using (I) steric bulk on the donor site, (II) alteration of the offset between the nitrogen donor groups, and (III) increase in the difference in the linker lengths for steering orientational self-sorting to the *HHTT*- Pd_2L_4 isomer and crystal structure of *HHTT*-**5.18**. (b) Illustration of heuristics D , the maximal sum of the deviations of the Pd(II)-nitrogen bonds from the average PdN_4 plane and q , the minimum square planar order parameter for estimating the stability of cage isomers.

Zhang, Wang and coworkers increased the asymmetry of ligand **5.11** by increasing the length of one linker (principle III). Self-assembly with phenyl spaced ligand **5.16** afforded *HHTT*- $\text{Pd}_2(5.16)_4$ as the main species along with traces of other species. A further increase in ligand dissymmetry by incorporating a phenyl-alkynyl-spaced linker in **5.17** afforded the same isomer *HHTT*- $\text{Pd}_2(5.17)_4$, **5.18** as the sole species. DFT computations support that by increasing the linker length, the energy difference between the *HHTT*- and the *HHHT*-isomer increases ($5.11 < 5.16 < 5.17$).¹⁸⁸

In a combined theoretical and experimental study, Jelfs, Lewis, and Tarzia developed a workflow that aims at predicting the selective orientational self-sorting to *HHTT*- Pd_2L_4 cages with high throughput. For ranking the stability of the isomers of a virtual library of asymmetric ligands, three heuristics were considered: the energy difference ΔE between the two most stable isomers, as well as the sum of the deviations of the Pd(II)-nitrogen bonds from the average PdN_4 plane D and the minimum square planar order parameter q (in an ideal square planar environment $D = 0$ and $q = 1$, Fig. 22b). From the



isomers that were predicted to have parameters favouring the formation of a single *HHTT*-isomer ($\Delta E \geq 6 \text{ kJ mol}^{-1}$, $D < 0.1 \text{ \AA}$, $q > 0.95$), three were already shown to selectively form the *HHTT*-isomer (ligands **5.13**, **5.14**, and **5.15**). For testing the fidelity of the approach, the authors experimentally investigated the self-assembly of five new ligands. While single isomers were obtained for ligands strictly obeying the set structural parameters, the approach was less reliable for small ΔE values. Hence, the parameters considered do not seem to describe all factors influencing the orientational self-sorting. Omitted effects include the templation by anions or solvents as well as ligand flexibility.¹⁸⁹

In 2019, Ogata and Yuasa studied the self-assembly of ligand **5.19** which is equipped with a *para*-pyridine as well as an imidazole donor moiety. When **5.19** and the Pd(II) source are combined in a 2 : 1 ratio, a single isomer Pd₂(**5.19**)₄, **5.20** was obtained (Fig. 23a). NMR spectroscopic investigations in combination with DFT computations support that the *HHTT*-isomer is favoured. Steric repulsion between pyridine and imidazole protons caused a twisting of the assembly which led to the existence of two diastereomers if chiral ligand **5.19b** was employed. The authors furthermore exploited the superior donor strength of imidazole: upon decreasing the **5.19**: Pd(II) stoichiometry, open structure Pd₁(**5.19**)₄ was obtained, where the ligand only coordinates *via* the imidazole end. The structural conversion was accompanied by the release of an encapsulated anion.¹⁹⁰

The group of Chand employed ligand **5.21** with 4-aniline and *meta*-pyridine ends. The self-assembly was studied with various Pd(II) precursors using DMSO as the solvent. In each case, formation of a single isomer assigned to *HHTT*-Pd₂(**5.21**)₄, **5.22** was achieved, noteworthy with the aniline amino substituents

acting as donors besides the pyridines (Fig. 23b). Furthermore, robust selectivity for this isomer was observed at different temperatures as well as at varying concentrations. The authors performed DFT computations and molecular dynamics (MD) studies which supported that the superior stability of the *HHTT*-isomer is driven on one hand by its least conformational strain and on the other hand by stabilization through electrostatic and van-der-Waals interactions.¹⁹¹

Additionally, the same group reported an asymmetric ligand with *meta*-pyridine and 3-aniline donor groups separated by an amide bond. This short ligand was shown to assemble to form a single isomer Pd₃L₆ which was, based on spectroscopic and computational investigations, assigned to the all-*HHTT* isomer.¹⁹²

Lewis presented in 2021 the selective formation of *HHTT*-Pd₂(**5.23**)₄, **5.24** isomers with ligands **5.23**, terminated with an isoquinoline and a triazole donor (Fig. 23c). The latter was introduced in the last step *via* a CuAAC click reaction, allowing for easy functionalization of the ligand. While the selectivity for the *HHTT*-isomer proved robust across a variety of different electronic situations for the triazole donor (ligands **5.23a-f**), variation of the donor strength affected the self-assembly when ligand and Pd(II) were combined in a 4 : 1 ratio. For phthalimide-appended ligand **5.23e**, a mononuclear Pd₁(**5.23e**)₄ complex was observed, similar to the work of Ogata and Yuasa.¹⁹⁰ The ligand coordinates here solely *via* the isoquinoline donor moieties as the substituent in **5.23e** (further) reduces the triazole donor strength. In contrast, when ligand **5.23a** was employed, a mixture of free ligand, mono-, and dinuclear species was observed. Exploiting the isostructural nature of ligands **5.23**, the author showed that statistical mixtures of homo- and heteroleptic cages were obtained when multiple ligands were combined.¹⁹³

Zhang and coworkers substituted the *meta*-pyridine in ligand **5.11** with a *para*-pyridine resulting in an increased binding angle of ligand **5.25** (Fig. 24a). In accordance with the trends observed for homoleptic cages (Chapter 2), this resulted in the formation of a species of higher nuclearity, here a three-membered ring. Noteworthy, geometric constraints prompted selective formation of all-*HHTT*-Pd₃(**5.25**)₆, isomer **5.26**.¹⁸⁸

Recently, Chand and coworkers reported the formation of a tetrahedral all-*HHTT*-Pd₄(**5.27**)₈, **5.28** by applying geometrically constrained ligand **5.27** with a bent angle of approximately 120° (Fig. 24b).¹⁹⁴

In 2021, Severin and coworkers reported tetrahedral and octahedral assemblies with asymmetric ligands **5.29** and **5.31** that possess wide bent angles (Fig. 24c and d). The number of possible isomers reaches 35 in the case of the tetrahedral assembly and to 112 in the case of the octahedral one. Astonishingly, the authors observed the formation of single isomers all-*HHTT*-Pd₄(**5.29**)₈, **5.30** and all-*HHTT*-Pd₆(**5.31**)₁₂, **5.32** in which two different ligand ends are arranged in a *cis*-configuration on each Pd(II) centre.¹⁹⁵ The preferred orientation of the ligands can be traced back to geometrical constraints, as was explored in a later study by the same authors on the basis of octahedral assemblies: ligand **5.31** can be characterized by the ratio of the two donor vectors *a/b*. Two ligands opposing each

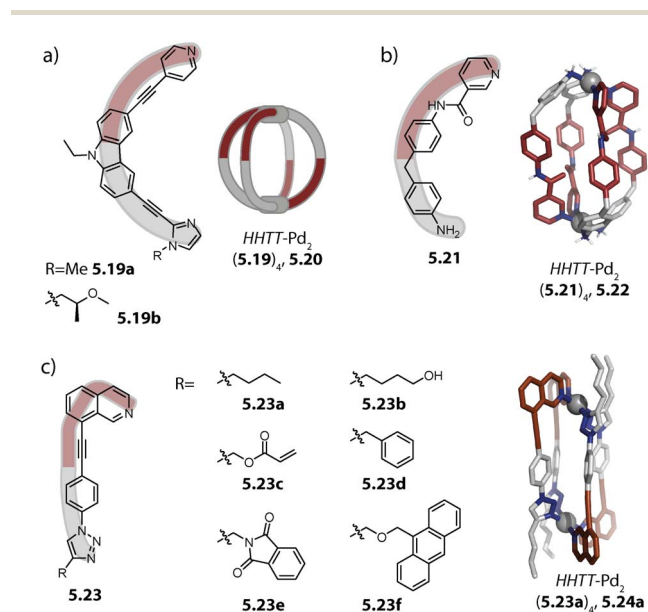


Fig. 23 Orientational self-sorting with ligands bearing differing donor moieties. (a) Ligand **5.19** combines pyridine and imidazole donors. (b) Ligand **5.21** comprises pyridine and aniline donors. (c) Ligand **5.23** has isoquinoline and triazole donors.



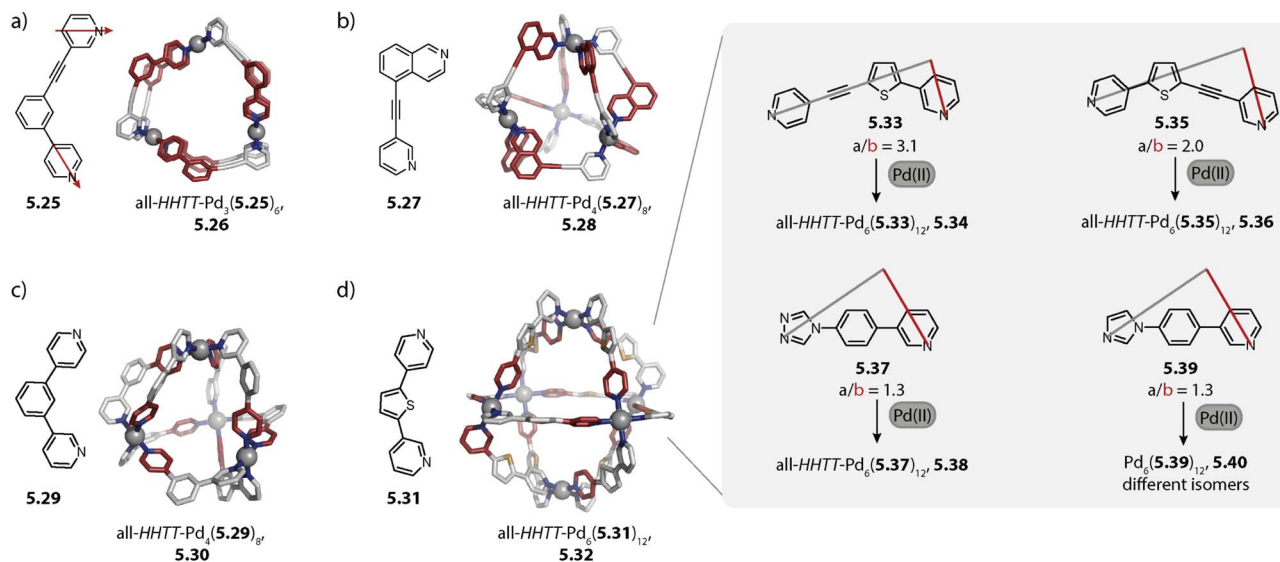


Fig. 24 Orientational self-sorting to the all-*HHTT* isomer in tri- (a), tetra- (b and c), and hexanuclear species (d). The inset shows the selectivity of the orientational self-sorting for ligands with differing ratios a/b .

other in the assembly can either be oriented *HH*, *TT*, or *HT*. Only the latter case allows for a Pd–Pd–Pd angle of 90° required for an octahedron. Conceivable limitations of the so-called *cis*-rule are a small a/b ratio as well as ligand flexibility. Both aspects were altered in alkyne-spaced ligands 5.33 and 5.35. Ligand 5.33, possessing a large a/b ratio cleanly assembles, similarly to ligand 5.31, to form octahedron all-*HHTT*-Pd₆(5.33)₁₂, 5.34. When ligand 5.35 with a decreased asymmetry was employed, prolonged heating was required for the selective formation of all-*HHTT*-Pd₆(5.35)₁₂, 5.36. The approximate rectangular bent angle that allows for the formation of an octahedron is maintained in ligands 5.37 and 5.39, possessing triazole or imidazole donors, respectively. The former strictly follows the *cis*-rule, yielding all-*HHTT*-Pd₆(5.37)₁₂, 5.38. In contrast, a mixture of isomers Pd₆(5.39)₁₂, 5.40 was obtained with ligand 5.39, that differs merely in its electronic situation. The selectivity might be hampered by the formation of kinetically trapped intermediates due to the higher basicity of the imidazole donor moiety, as hypothesized by the authors.¹⁹⁶

Only very recently, Chand and coworkers were able to combine the non-statistical self-assembly of heteroleptic cages with orientational self-sorting. For this, they employed two strategies: the SCA approach for forming *cis*-Pd₂A₂B₂ cages as well as geometrical constraints for driving orientational self-sorting. For the former, the two ligands 5.27 and 5.41 were designed in a way that their donor vectors are convergent and divergent, respectively (for ligand 5.41 in both conformations, Fig. 25a). The geometric complementarity of 5.27 and 5.41 reduced the number of possible isomers from 31 to six. Furthermore, the structural constraints were expected to only allow a ligand arrangement in which the amido-pyridine ends of 5.41 oppose the pyridine ends of ligands 5.27 and the pyridine ends of 5.41 oppose the isoquinoline ends of 5.27. This leaves two possible isomers, namely *HHH'H'*-Pd₂(5.27)₂(5.41)₂ and *HTH'T'*-Pd₂(5.27)₂(5.41)₂. The authors found that *HTH'T'*-

Pd₂(5.27)₂(5.41)₂, 5.42 was formed selectively (Fig. 25b). The preference for the antiparallel arrangement of the ligands can be explained with increased steric strain when the two isoquinoline donors of 5.27 sit in *cis*-position to each other. Interestingly, substitution of the alkyne spacer in 5.27 by a phenyl spacer hampered self-sorting to a defined species. It was hypothesized that this is due to a length mismatch.¹⁹⁴ Lewis and coworkers, on the other hand, achieved isomer-selective assembly of low-symmetry heteroleptic Pd₂LA₂LB₂ cages through a combination of shape-complementary assembly and coordination sphere engineering.¹⁹⁷

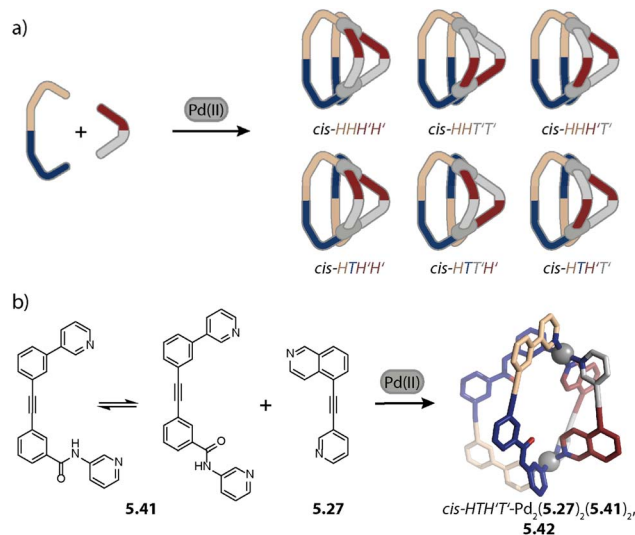


Fig. 25 (a) Possible isomers for the self-assembly of two asymmetric, shape-complementary ligands in the absence of factors driving orientational self-sorting (enantiomers are not shown). (b) Selective assembly of *cis*-*HTH'T'*-5.42 through the SCA approach in combination with geometrical constraints driving orientational self-sorting.



The repertoire of self-assemblies with asymmetric ligands has greatly expanded within recent years. In most cases, ligands with roughly parallel binding vectors have been employed for obtaining dinuclear Pd_2L_4 cages, which possess the smallest number of possible isomers. Additionally, all-*HHTT* isomers of higher nuclearity species have been reported, that are three-membered rings, tetrahedra, and octahedra. The development of strategies for orientational self-sorting was highly shaped by already existing ones for the non-statistical self-sorting of heteroleptic cages, that are strategies exploiting steric or geometric constraints or a combination thereof. This showcased that the design principles are transferrable; hence, it is expected that the two areas will mutually benefit from each other in the future. Furthermore, orientational self-sorting allows for differentiable coordination environments; however, the number of examples for species dealing with the consequences of this phenomenon remains scarce. Recently, strategies for integrative and orientational self-sorting have been transferred to $\text{Pt}(\text{II})$ cage chemistry by Preston, using a specific $\text{Pt}(\text{II})$ precursor that ensures sufficient kinetic accessibility of the thermodynamic products in the assembly reaction.¹⁹⁸

6. Chiral cages

Chirality in metallo-supramolecular assemblies can be a consequence of the spatial arrangement of the components upon assembly formation or can stem from the inherent chirality of one or more components. The former, commonly referred to as a “soft” approach is promising for stimuli responsive systems due to its dynamicity. Chiral interaction partners, such as guests, can then transfer their chirality on the assembly *via* non-covalent interactions. The latter, “hard” approach can be advantageous for functional materials, *e.g.* for chiral sensing or asymmetric catalysis, due to its robustness.¹⁹⁹ Employing chiral ligands as a racemic mixture endows the self-assembly process with a further layer of complexity: in the absence of self-sorting, a mixture of different diastereoisomers can be obtained. Narcissistic chiral self-sorting occurs under self-recognition, leading to a mixture of two assemblies bearing only ligands of

the same handedness. In contrast, social chiral self-sorting is based on self-discrimination and may lead to overall achiral *meso*-compounds.

The groups of Kubik and Lützen as well as our group used chiral, peptide-based ligands for assembling three-membered rings²⁰⁰ and doubly interlocked lemniscates.⁹⁸ Furthermore, Lützen and coworkers reported homochiral tetrahedra, cubes, and dodecanuclear spheres based on BINOL ligands.^{124,127} Liu and coworkers utilized the BINOL-based cube for obtaining CPL-active host-guest complexes with BODIPY guests.²⁰¹ The group of Zysman-Colman formed three- and four-membered rings for photophysical investigations based on paracyclophane²⁰² and homochiral $\text{Ir}(\text{III})$ ²⁰³ complexes. Furthermore, Rissanen and coworkers reported a series of assemblies with ligands based on bile acids,^{184,185} as described in Chapter 5. Aside from these architectures that were assembled from enantiomerically pure ligands, the chiral self-sorting of racemic mixtures of ligands gained particular attention and will be the focus of this chapter.

Lützen and coworkers reported in 2013 on the first system showing narcissistic chiral self-sorting in Pd_2L_4 cages. The authors studied the self-assembly of BINOL-based ligands *M*- or *P*-**6.1** possessing axial chirality. When **6.1** is employed as a racemic mixture, six different isomers can potentially be obtained: two homochiral cages, two achiral, diastereomeric cages (*meso-cis* and *meso-trans*), and two cages carrying both enantiomers in a 3 : 1 ratio. Self-assembly occurred under chiral self-recognition, leading to a narcissistic mixture of the two homochiral cages $\text{Pd}_2(\text{M-6.1})_4$, *MMMM-6.2* and $\text{Pd}_2(\text{P-6.1})_4$, *PPPP-6.2* (Fig. 26a).²⁰⁴

Recently, we reported the narcissistic chiral self-sorting of Tröger-based ligands *M*-**L** and *P*-**L** into $\text{Pd}_2(\text{M-L})_4$ and $\text{Pd}_2(\text{P-L})_4$ capsules upon encapsulation of fullerene guests or upon crystallization, respectively.⁷⁸

Social chiral self-sorting in Pd_2L_4 cages, on the other hand, was reported by us in 2019 for the first time. Ligands *M*- and *P*-**6.3**, based on a helicene backbone, assemble with $\text{Pd}(\text{II})$ to form dinuclear homochiral cages $\text{Pd}_2(\text{M/P-6.3})_4$, *MMMM/PPPP-6.4*. When a racemic mixture of the ligand is employed, self-

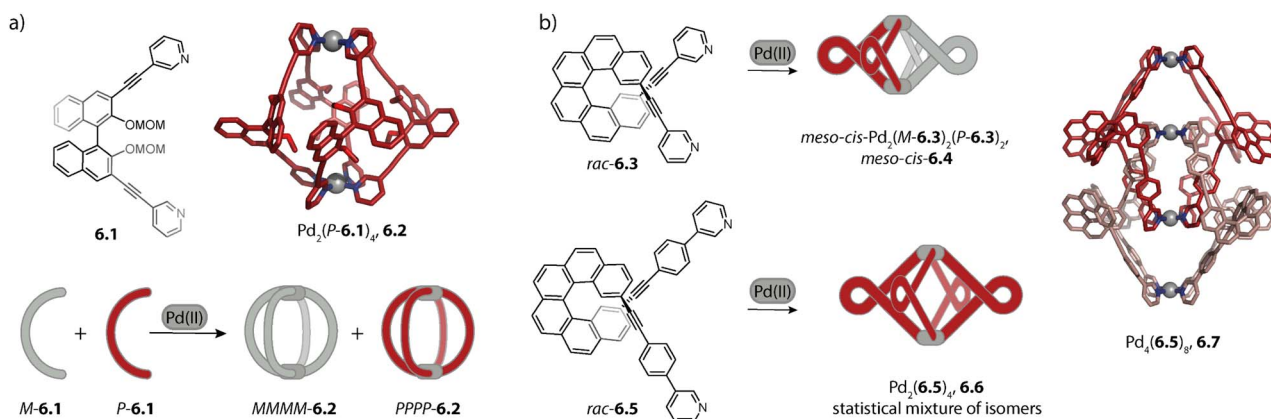


Fig. 26 (a) Narcissistic chiral self-sorting to a pair of enantiomeric lantern-shaped cages **6.2**. (b) Social chiral self-sorting to *meso-cis-6.4* that is lost upon prolongation of the linker.



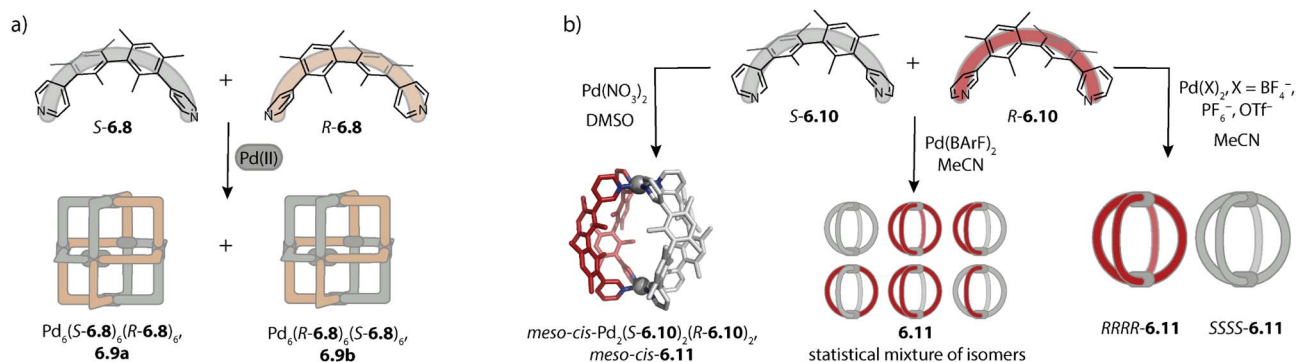


Fig. 27 (a) Social chiral self-sorting to a pair of cube enantiomers **6.9**. (b) Solvent- and counter anion-dependent chiral self-sorting with a racemic mixture of ligands **6.10**.

discrimination to *meso-cis*- $\text{Pd}_2(\text{M-6.3})_2(\text{P-6.3})_2$ and *meso-cis-6.4* is observed (Fig. 26b). Interestingly, a breakdown of chiral self-sorting was observed for ligands *M*- and *P*-**6.5**, differing only in the elongated linker between the helicene core and the pyridine donors. This was thought to originate from the increased distance between the chiral backbones in the assembly. The enantiopure monomeric cages $\text{Pd}_2(\text{M/P-6.5})_4$ and *MMMM/PPPP-6.6* were shown to discriminate between guests of opposite chirality. Furthermore, in line with the criteria for the formation of interlocked double cages (Chapter 2), elongated ligand **6.5** assembled in acetonitrile to form double cage $\text{Pd}_4(\text{6.5})_8$, **6.7**.¹³⁴

Recently, the group of Natarajan observed social chiral self-sorting in Pd_6L_{12} cubes. Ligands **6.8** possess axial chirality and a rectangular bent angle. Due to the increased number of components in hexanuclear in comparison to dinuclear cages, the number of possible isomers reaches 145. Excitingly, the authors observed the selective formation of a racemic pair of heterochiral cages $\text{Pd}_6(\text{S-6.8})_6(\text{R-6.8})_6$, **6.9** (Fig. 27a).²⁰⁵

The same group additionally explored the chiral self-sorting with an isomer of **6.8**, ligand **6.10** possessing *meta*-pyridine donors. Intriguingly, when BARF^- was employed as the counter

anion which, owing to its large size, resides outside of the narrow cavity, no chiral self-sorting was observed (Fig. 27b). However, social chiral self-sorting to *meso-cis*- $\text{Pd}_2(\text{S-6.10})_2(\text{R-6.10})_2$, *meso-cis-6.11* was observed in the presence of NO_3^- in DMSO and narcissistic chiral self-sorting to the homochiral pair *SSSS-6.11* and *RRRR-6.11* with BF_4^- , PF_6^- and OTf^- . Chiral self-sorting was furthermore shown to be solvent-dependent for NO_3^- or BF_4^- as counter anions: with NO_3^- , a solvent switch from DMSO to acetonitrile resulted in a mixture of *meso-cis-6.11* and a pair of homochiral cages. Additionally, a mixture of different isomers was obtained for the self-assembly with BF_4^- in DMSO.²⁰⁶

We recently carried out a thorough investigation of the solvent dependency of chiral self-sorting in helical cages with ligands **6.12** and **6.14** that are based on axially chiral 1,1'-biazulene-2,2'-diamine, a structural motif reminiscent of widely used BINOL (with azulene being an isomer of the therein contained naphthalene). Short ligand **6.12** and long ligand **6.14** both assemble to form Pd_2L_4 cages; however, a striking difference in chiral self-sorting was observed. In DMSO and acetonitrile, racemic **6.12** selectively forms *meso-trans*- $\text{Pd}_2(\text{S-6.12})_2(\text{R-6.12})_2$ (**6.13**)

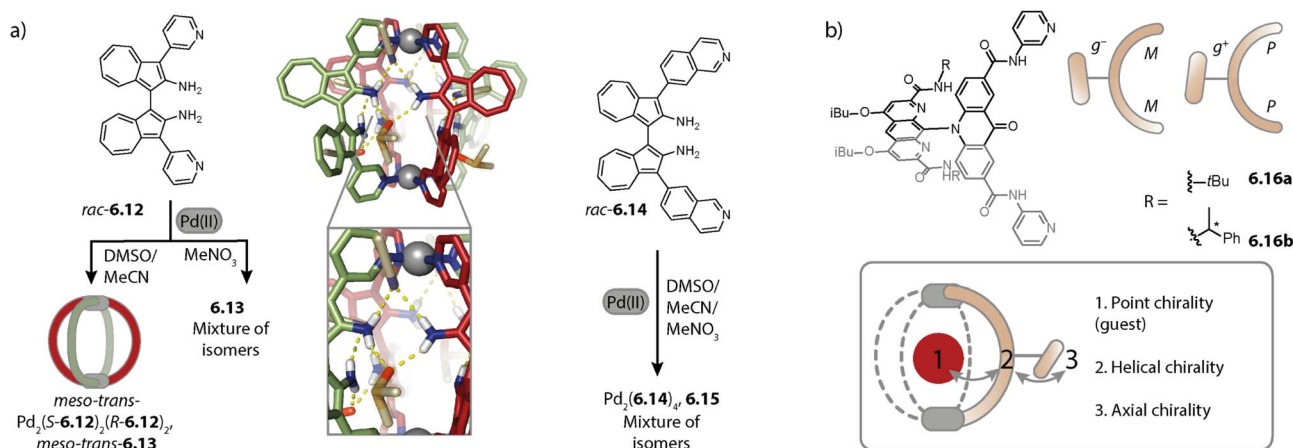


Fig. 28 (a) Solvent dependent chiral social self-sorting with biazulene-based ligand **6.12**; the expansion of the X-ray structure shows solvent molecules acting as tethers between two adjacent ligands. (b) Bidirectional transmission of chirality between the cavity and the periphery of the cage.



6.12)₂, *meso-trans*-**6.13** while a mixture of isomers Pd₂(**6.14**)₄, **6.15** is obtained with racemic **6.14** under the same conditions (Fig. 28a). Close examination of the crystal structure of *meso-trans*-**6.13** in combination with DFT computations including explicit solvation reveals that the *meso-trans* isomer is stabilized by individual solvent molecules that connect two adjacent ligands *via* hydrogen bonding. Accordingly, changing the solvent to nitromethane, a significantly weaker H-bond acceptor, resulted in a loss of chiral self-sorting (to be gained back by adding small amounts of DMSO). Such an H-bond tethering is not possible in **6.15** due to the larger distances between the amino groups in the ligand backbones.

Examples for the transfer of chirality from homochiral Pd_nL_{2n} cages to encapsulated guests were reported, either giving rise to chiral guest discrimination¹³⁴ or to induced guest chirality.^{77,78,201} Conversely, guest-to-host chirality transfer has been observed in terms of the stabilization of chirality in dynamically chiral hosts through binding of enantiopure guests.^{38,46,89}

Very recently, Gan and coworkers reported a bidirectional transmission of chiral information between a guest and the exterior of a helical cage. The authors incorporated in helicates Pd₂(**6.16**)₄ another stereogenic centre: the 1,8-diazaanthracene segments can rotate relative to the ligand backbone, giving rise to different *gauche* conformations (Fig. 28b). Owing to steric constraints, helical and axial chirality are linked; more precisely, a *PP*-conformation of the ligand favours a positive *gauche* conformation (*g*⁺) while a *MM*-conformation favours *g*⁻. Achiral ligand **6.16a** assembles to form *meso-cis*-Pd₂(*MM*-**6.16a**)₂(*PP*-**6.16a**)₂ incorporating both *gauche* conformations. Encapsulation of L- or D-tartrate guests induced a preference for one of the *gauche* conformations, meaning that the chirality is transmitted to the cage periphery. Homo-axial chirality could also be achieved through incorporation of point chirality in the appended segment of ligand **6.16b**. Chiral helicates Pd₂(*MM*-**6.16b**)₄ and Pd₂(*PP*-**6.16b**)₄ were shown to discriminate between L- and D-tartrate guests.²⁰⁷

Natural or synthetic chiral backbones have been employed for bestowing coordination assemblies of various nuclearities with chirality. We paid particular attention to the chiral self-sorting of racemic mixtures of ligands to either *meso*-cages or a mixture of the two enantiopure cages. Prolongation of the linker can result in a loss of chiral self-sorting due to the increased distance between the ligand backbones. Furthermore, the solvent and the counter anions can have a pivotal effect on the outcome. A possible origin for the former is the specific interactions between solvent molecules and the ligands at the portals of the cage. We also touched upon a recent example in which the soft and hard chirality of helicates was exploited for chirality transmission from or to an encapsulated guest, respectively.

7. Conclusion

To conclude, we surveyed the topological diversity of di- to octanuclear assemblies that are accessible upon simple combination of bis-monodentate organic ligands and Pd(II)

cations. The structure of the organic ligand, *i.e.* its geometry and steric demand, plays a key role in the self-assembly outcome. The stiffness of small, aromatic ligands allows them to be simplified as simple geometric units which facilitate structure prediction. In contrast, solvent and counter anions often exert a templation effect on self-assembly with highly flexible ligands, rendering their topologies notoriously more difficult to predict. Long organic ligands allow for mechanical interlocking which can be inhibited by equipping the ligands with endohedral steric bulk. Strategies for increasing structural complexity include the formation of heteroleptic cages with multiple organic building blocks as well as the usage of asymmetric ligands. For steering the self-sorting towards a defined species, the ligands have been endowed with geometric constraints, steric bulk has been installed at the backbone or at the donor moieties, and attractive as well as repulsive inter-ligand interactions have been exploited. Concerning dinuclear heteroleptic cages, the maximum level of self-sorting to provide Pd₂ABCD cages with four differentiable ligands has been mastered very recently. Several larger heteroleptic architectures have been reported, most of them carrying two different kinds of ligands. Orientational self-sorting with asymmetric ligands flourished in the last couple of years and defined architectures with up to six Pd(II) nodes have been attained. We have also treated self-sorting effects with chiral ligands; chiral narcissistic and social self-sorting was shown to be highly dependent on solvent and counter anions.

We anticipate that an enhanced understanding of the factors governing the formation of specific topologies as well as the reliability of strategies for selective self-sorting will allow the field to shift progressively towards more and more complex functional nanosystems. On one hand, the assemblies can be harnessed as platforms for combining multiple functional moieties in a combinatorial fashion. The overall assembly properties emerge then from the interplay of the various components, which holds promise for highly selective molecular receptors, catalysts, and drug delivery systems. Another field of application may be energy harvesting materials, *i.e.* in terms of gaining supramolecular control over the morphology of donor/acceptor-based charge separating layers in photovoltaic devices. While examples for functional cages incorporating two kinds of interoperating ligands have been reported (many by our lab), the newly developed strategies for the non-statistical combination of more than two kinds of ligands have not yet been exploited for applicable cages. On the other hand, as an additional layer of complexity, populations of coexisting assemblies have been reported by us very recently. In biological systems, numerous processes occur simultaneously, either in an orthogonal fashion or are closely interconnected. The dynamic nature of the coordination bond as well as the possibility of incorporating responsive moieties can be exploited for creating networks of communicating and coexisting assemblies. This offers great potential for the transmission and transduction of information or energy, which is of interest for the development of smart materials and of the field of systems chemistry.



Data availability

No primary research results, software or code have been included and no new data were generated or analysed as part of this review.

Author contributions

L. N. compiled the draft and created the figures. G. H. C. and L. N. both discussed and revised the manuscript.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was financed by the Deutsche Forschungsgemeinschaft (DFG, German Research Foundation) under GRK2376, project number 331085229 ("Confinement-Controlled Chemistry") and Germany's Excellence Strategy EXC2033, project number 390677874 ("RESOLV").

Notes and references

- J.-M. Lehn, *Science*, 1993, **260**, 1762.
- A. Stank, D. B. Kokh, J. C. Fuller and R. C. Wade, *Acc. Chem. Res.*, 2016, **49**, 809.
- K. Selvakumar, L. Motiei and D. Margulies, *J. Am. Chem. Soc.*, 2015, **137**, 4892.
- P. Parkinson, C. E. I. Knappe, N. Kamonsutthipajit, K. Sirithip, J. D. Matchak, H. L. Anderson and L. M. Herz, *J. Am. Chem. Soc.*, 2014, **136**, 8217.
- G. M. Whitesides and B. Grzybowski, *Science*, 2002, **295**, 2418.
- M. E. Briggs and A. I. Cooper, *Chem. Mater.*, 2017, **29**, 149.
- F. Beuerle and B. Gole, *Angew. Chem., Int. Ed.*, 2018, **57**, 4850.
- M. Mastalerz, *Angew. Chem., Int. Ed.*, 2010, **49**, 5042.
- T. Kunde, T. Pausch and B. M. Schmidt, *Eur. J. Org. Chem.*, 2021, **2021**, 5844.
- L. R. MacGillivray and J. L. Atwood, *Nature*, 1997, **389**, 469.
- J. Rebek, *Acc. Chem. Res.*, 2009, **42**, 1660.
- S. K. Yang and S. C. Zimmerman, *Isr. J. Chem.*, 2013, **53**, 511.
- L. J. Prins, D. N. Reinhoudt and P. Timmerman, *Angew. Chem., Int. Ed.*, 2001, **40**, 2382.
- M. M. J. Smulders, I. A. Riddell, C. Browne and J. R. Nitschke, *Chem. Soc. Rev.*, 2012, **42**, 1728.
- T. R. Cook and P. J. Stang, *Chem. Rev.*, 2015, **115**, 7001.
- M. Han, D. M. Engelhard and G. H. Clever, *Chem. Soc. Rev.*, 2014, **43**, 1848.
- M. Pan, K. Wu, J.-H. Zhang and C.-Y. Su, *Coord. Chem. Rev.*, 2019, **378**, 333.
- F. Würthner, C.-C. You and C. R. Saha-Möller, *Chem. Soc. Rev.*, 2004, **33**, 133.
- C. M. Hong, R. G. Bergman, K. N. Raymond and F. D. Toste, *Acc. Chem. Res.*, 2018, **51**, 2447.
- M. Fujita, M. Tominaga, A. Hori and B. Therrien, *Acc. Chem. Res.*, 2005, **38**, 369.
- D. Bardhan and D. K. Chand, *Chem.–Eur. J.*, 2019, **25**, 12241.
- S. De, K. Mahata and M. Schmittel, *Chem. Soc. Rev.*, 2010, **39**, 1555.
- R. W. Saalfrank, A. Stark, K. Peters and H. G. von Schnering, *Angew. Chem., Int. Ed.*, 1988, **27**, 851.
- B. Hasenknopf, J.-M. Lehn, N. Boumediene, A. Dupont-Gervais, A. V. Dorsselaer, B. Kneisel and D. Fenske, *J. Am. Chem. Soc.*, 1997, **119**, 10956.
- D. L. Caulder, R. E. Powers, T. N. Parac and K. N. Raymond, *Angew. Chem., Int. Ed.*, 1998, **37**, 1840.
- P. J. Stang, *Chem.–Eur. J.*, 1998, **4**, 19.
- M. Fujita, J. Yazaki and K. Ogura, *J. Am. Chem. Soc.*, 1990, **112**, 5645.
- D. A. McMorran and P. J. Steel, *Angew. Chem., Int. Ed.*, 1998, **37**, 3295.
- Z. T. Avery, J. L. Algar and D. Preston, *Trends Chem.*, 2024, **6**, 352.
- J. E. M. Lewis and J. D. Crowley, *ChemPlusChem*, 2020, **85**, 815.
- D. Bokotial, K. Acharyya, A. Chowdhury and P. S. Mukherjee, *Angew. Chem., Int. Ed.*, 2024, **63**, e202401136.
- T. K. Piskorz, V. Martí-Centelles, R. L. Spicer, F. Duarte and P. J. Lusby, *Chem. Sci.*, 2023, **14**, 11300.
- V. Martí-Centelles, A. L. Lawrence and P. J. Lusby, *J. Am. Chem. Soc.*, 2018, **140**, jacs.7b12146.
- R. L. Spicer, A. D. Stergiou, T. A. Young, F. Duarte, M. D. Symes and P. J. Lusby, *J. Am. Chem. Soc.*, 2020, **142**, 2134.
- T. A. Young, V. Martí-Centelles, J. Wang, P. J. Lusby and F. Duarte, *J. Am. Chem. Soc.*, 2020, **142**, 1300.
- S. Pradhan, A. Bhattacharyya and R. P. John, *Tetrahedron Lett.*, 2016, **57**, 1532.
- M. Xu, B. Sun, D. A. Poole, E. O. Bobylev, X. Jing, J. Wu, C. He, C. Duan and J. N. H. Reek, *Chem. Sci.*, 2023, **14**, 11699.
- I. Regeni, B. Chen, M. Frank, A. Baksi, J. J. Holstein and G. H. Clever, *Angew. Chem., Int. Ed.*, 2021, **60**, 5673.
- A. B. S. Elliott, J. E. M. Lewis, H. van der Salm, C. J. McAdam, J. D. Crowley and K. C. Gordon, *Inorg. Chem.*, 2016, **55**, 3440.
- B. Aikman, R. Bonsignore, B. Woods, D. Doellerer, R. Scotti, C. Schmidt, A. A. Heidecker, A. Pöthig, E. J. Sayers, A. T. Jones and A. Casini, *Dalton Trans.*, 2022, **51**, 7476.
- K. E. Ebbert, E. Benchimol, A. Platzek, C. Drechsler, J. Openy, S. Hasegawa, J. J. Holstein and G. H. Clever, *Angew. Chem., Int. Ed.*, 2024, **63**, e202413323.
- C. Li, B. Zhang, Y. Dong, Y. Li, P. Wang, Y. Yu, L. Cheng and L. Cao, *Dalton Trans.*, 2020, **49**, 8051.
- S. Ganta, J.-H. Borter, C. Drechsler, J. J. Holstein, D. Schwarzer and G. H. Clever, *Org. Chem. Front.*, 2022, **9**, 5485.
- V. Croué, S. Krykun, M. Allain, Y. Morille, F. Aubriet, V. Carré, Z. Voitenko, S. Goeb and M. Sallé, *New J. Chem.*, 2017, **41**, 3238.



- 45 R. A. S. Vasdev, J. A. Findlay, A. L. Garden and J. D. Crowley, *Chem. Commun.*, 2019, **55**, 7506.
- 46 R.-J. Li, J. J. Holstein, W. G. Hiller, J. Andréasson and G. H. Clever, *J. Am. Chem. Soc.*, 2019, **141**, 2097.
- 47 R. Li, M. Han, J. Tessarolo, J. J. Holstein, J. Lübben, B. Dittrich, C. Volkmann, M. Finze, C. Jenne and G. H. Clever, *ChemPhotoChem*, 2019, **3**, 378.
- 48 S. Fu, Q. Luo, M. Zang, J. Tian, Z. Zhang, M. Zeng, Y. Ji, J. Xu and J. Liu, *Mater. Chem. Front.*, 2019, **3**, 1238.
- 49 J. Zhu, X. Chen, X. Jin and Q. Wang, *Chin. Chem. Lett.*, 2023, **34**, 108002.
- 50 R. G. DiNardi, A. O. Douglas, R. Tian, J. R. Price, M. Tajik, W. A. Donald and J. E. Beves, *Angew. Chem., Int. Ed.*, 2022, **61**, e202205701.
- 51 S. Ganta, C. Drechsler, Y. Chen and G. H. Clever, *Chem.–Eur. J.*, 2022, **28**, e202104228.
- 52 S. Saha, Y.-T. Chen, S. Ganta, M. Gilles, B. Holzapfel, P. Lill, H. Rehage, C. Gatsogiannis and G. H. Clever, *Chem.–Eur. J.*, 2022, **28**, e202103406.
- 53 S. Saha, B. Holzapfel, Y.-T. Chen, K. Terlinden, P. Lill, C. Gatsogiannis, H. Rehage and G. H. Clever, *J. Am. Chem. Soc.*, 2018, **140**, 17384.
- 54 M. D. Ward, *Chem. Commun.*, 2024, **60**, 10464.
- 55 Y. Xue, X. Hang, J. Ding, B. Li, R. Zhu, H. Pang and Q. Xu, *Coord. Chem. Rev.*, 2020, **430**, 213656.
- 56 H. Takezawa and M. Fujita, *Bull. Chem. Soc. Jpn.*, 2021, **94**, 2351.
- 57 W.-L. Jiang, B. Huang, X.-L. Zhao, X. Shi and H.-B. Yang, *Chem*, 2023, **9**, 2655.
- 58 S. Hasegawa, A. Baksi, B. Chen and G. H. Clever, *Org. Mater.*, 2022, **4**, 222.
- 59 M. Yoshizawa and L. Catti, *Acc. Chem. Res.*, 2019, **52**, 2392.
- 60 S. Yadav, P. Kannan and G. Qiu, *Org. Chem. Front.*, 2020, **7**, 2842.
- 61 D. Fujita, Y. Ueda, S. Sato, N. Mizuno, T. Kumasaka and M. Fujita, *Nature*, 2016, **540**, 563.
- 62 D. Fujita, Y. Ueda, S. Sato, H. Yokoyama, N. Mizuno, T. Kumasaka and M. Fujita, *Chem*, 2016, **1**, 91.
- 63 R. Chakrabarty, P. S. Mukherjee and P. J. Stang, *Chem. Rev.*, 2011, **111**, 6810.
- 64 S. Pullen, J. Tessarolo and G. H. Clever, *Chem. Sci.*, 2021, **12**, 7269.
- 65 L. K. Moree, L. A. V. Faulkner and J. D. Crowley, *Chem. Soc. Rev.*, 2024, **53**, 25.
- 66 K. Acharyya, S. Mukherjee and P. S. Mukherjee, *J. Am. Chem. Soc.*, 2013, **135**, 554.
- 67 D. Preston, K. M. Patil, A. T. O'Neil, R. A. S. Vasdev, J. A. Kitchen and P. E. Kruger, *Inorg. Chem. Front.*, 2020, **7**, 2990.
- 68 S. M. Jansze, M. D. Wise, A. V. Vologzhanina, R. Scopelliti and K. Severin, *Chem. Sci.*, 2017, **8**, 1901.
- 69 N. Kishi, Z. Li, K. Yoza, M. Akita and M. Yoshizawa, *J. Am. Chem. Soc.*, 2011, **133**, 11438.
- 70 M. Yamashina, T. Yuki, Y. Sei, M. Akita and M. Yoshizawa, *Chem.–Eur. J.*, 2015, **21**, 4200.
- 71 G. H. Clever, W. Kawamura, S. Tashiro, M. Shiro and M. Shionoya, *Angew. Chem., Int. Ed.*, 2012, **51**, 2606.
- 72 N. Kishi, Z. Li, Y. Sei, M. Akita, K. Yoza, J. S. Siegel and M. Yoshizawa, *Chem.–Eur. J.*, 2013, **19**, 6313.
- 73 P. Liao, B. W. Langloss, A. M. Johnson, E. R. Knudsen, F. S. Tham, R. R. Julian and R. J. Hooley, *Chem. Commun.*, 2010, **46**, 4932.
- 74 D. P. August, G. S. Nichol and P. J. Lusby, *Angew. Chem., Int. Ed.*, 2016, **55**, 15022.
- 75 S. Hasegawa, S. L. Meichsner, J. J. Holstein, A. Baksi, M. Kasanmascheff and G. H. Clever, *J. Am. Chem. Soc.*, 2021, **143**, 9718.
- 76 B. Chen, J. J. Holstein, S. Horiuchi, W. G. Hiller and G. H. Clever, *J. Am. Chem. Soc.*, 2019, **141**, 8907.
- 77 H. Sasafuchi, M. Ueda, N. Kishida, T. Sawada, S. Suzuki, Y. Imai and M. Yoshizawa, *Chem*, 2024, 102332.
- 78 E. Benchimol, H. M. O'Connor, B. Schmidt, N. Bogo, J. J. Holstein, J. I. Lovitt, S. Shanmugaraju, C. J. Stein, T. Gunnlaugsson and G. H. Clever, *Angew. Chem., Int. Ed.*, 2025, **64**, e202421137.
- 79 A. Walther, I. Regeni, J. J. Holstein and G. H. Clever, *J. Am. Chem. Soc.*, 2023, **145**, 25365.
- 80 W. M. Bloch, Y. Abe, J. J. Holstein, C. M. Wandtke, B. Dittrich and G. H. Clever, *J. Am. Chem. Soc.*, 2016, **138**, 13750.
- 81 L.-P. Zhou and Q.-F. Sun, *Chem. Commun.*, 2015, **51**, 16767.
- 82 B. Chen, S. Horiuchi, J. J. Holstein, J. Tessarolo and G. H. Clever, *Chem.–Eur. J.*, 2019, **25**, 14921.
- 83 R. A. S. Vasdev, L. F. Gaudin, D. Preston, J. P. Jogy, G. I. Giles and J. D. Crowley, *Front. Chem.*, 2018, **6**, 563.
- 84 S. Sudan, F. Fadaei-Tirani and K. Severin, *Chem. Commun.*, 2023, **59**, 8258.
- 85 P. J. Steel and D. A. McMorran, *Chem.–Asian J.*, 2019, **14**, 1098.
- 86 H. S. Sahoo and D. K. Chand, *Dalton Trans.*, 2010, **39**, 7223.
- 87 D. Tripathy, A. K. Pal, G. S. Hanan and D. K. Chand, *Dalton Trans.*, 2012, **41**, 11273.
- 88 N. L. S. Yue, D. J. Eisler, M. C. Jennings and R. J. Puddephatt, *Inorg. Chem.*, 2004, **43**, 7671.
- 89 Q. Lin, L. Gao, B. Kauffmann, J. Zhang, C. Ma, D. Luo and Q. Gan, *Chem. Commun.*, 2018, **54**, 13447.
- 90 M. Sarkar, H. Dasary and D. K. Chand, *J. Organomet. Chem.*, 2021, **950**, 121984.
- 91 J. D. Crowley and E. L. Gavey, *Dalton Trans.*, 2010, **39**, 4035.
- 92 S. Ø. Scott, E. L. Gavey, S. J. Lind, K. C. Gordon and J. D. Crowley, *Dalton Trans.*, 2011, **40**, 12117.
- 93 S. M. McNeill, D. Preston, J. E. M. Lewis, A. Robert, K. Knerr-Rupp, D. O. Graham, J. R. Wright, G. I. Giles and J. D. Crowley, *Dalton Trans.*, 2015, **44**, 11129.
- 94 E. Benchimol, K. E. Ebbert, A. Walther, J. J. Holstein and G. H. Clever, *Chem.–Eur. J.*, 2024, **30**, e202401850.
- 95 S. Löffler, J. Lübben, L. Krause, D. Stalke, B. Dittrich and G. H. Clever, *J. Am. Chem. Soc.*, 2015, **137**, 1060.
- 96 W. M. Bloch, S. Horiuchi, J. J. Holstein, C. Drechsler, A. Wuttke, W. Hiller, R. A. Mata and G. H. Clever, *Chem. Sci.*, 2023, **14**, 1524.
- 97 D. Preston, S. M. McNeill, J. E. M. Lewis, G. I. Giles and J. D. Crowley, *Dalton Trans.*, 2016, **45**, 8050.



- 98 T. R. Schulte, J. J. Holstein, L. Schneider, A. Adam, G. Haberhauer and G. H. Clever, *Angew. Chem., Int. Ed.*, 2020, **59**, 22489.
- 99 R. Zhu, W. M. Bloch, J. J. Holstein, S. Mandal, L. V. Schäfer and G. H. Clever, *Chem.–Eur. J.*, 2018, **24**, 12976.
- 100 F. A. Pereira, T. Fallows, M. Frank, A. Chen and G. H. Clever, *Z. Anorg. Allg. Chem.*, 2013, **639**, 1598.
- 101 Q. Zhang, L. He, J.-M. Liu, W. Wang, J. Zhang and C.-Y. Su, *Dalton Trans.*, 2010, **39**, 11171.
- 102 A. Sautter, D. G. Schmid, G. Jung and F. Würthner, *J. Am. Chem. Soc.*, 2001, **123**, 5424.
- 103 W. M. Bloch, J. J. Holstein, W. Hiller and G. H. Clever, *Angew. Chem., Int. Ed.*, 2017, **56**, 8285.
- 104 I. Regeni, R. Chowdhury, K. Terlinden, S. Horiuchi, J. J. Holstein, S. Feldmann and G. H. Clever, *Angew. Chem., Int. Ed.*, 2023, **62**, e202308288.
- 105 K. Harris, D. Fujita and M. Fujita, *Chem. Commun.*, 2013, **49**, 6703–6712.
- 106 D. K. Chand, K. Biradha, M. Kawano, S. Sakamoto, K. Yamaguchi and M. Fujita, *Chem.–Asian J.*, 2006, **1**, 82.
- 107 K. Suzuki, M. Kawano and M. Fujita, *Angew. Chem., Int. Ed.*, 2007, **46**, 2819.
- 108 D. K. Chand, M. Fujita, K. Biradha, S. Sakamoto and K. Yamaguchi, *Dalton Trans.*, 2003, 2750.
- 109 X.-C. Zhou, Y.-Z. Tan, Y.-Y. Ge, Y.-L. Lai, D. Luo, X.-P. Zhou and D. Li, *Sci. China Chem.*, 2025, **1**, 68.
- 110 Y.-Y. Ge, X.-C. Zhou, J. Zheng, J. Luo, Y.-L. Lai, J. Su, H.-J. Zhang, X.-P. Zhou and D. Li, *Inorg. Chem.*, 2023, **62**, 4048.
- 111 R. A. S. Vasdev, J. A. Findlay, D. R. Turner and J. D. Crowley, *Chem.–Asian J.*, 2021, **16**, 39.
- 112 L. Neukirch, M. D. Kulas, J. J. Holstein and G. H. Clever, *Chem.–Eur. J.*, 2024, **30**, e202400132.
- 113 D. Preston, J. J. Sutton, K. C. Gordon and J. D. Crowley, *Angew. Chem., Int. Ed.*, 2018, **57**, 8659.
- 114 D. Samanta and P. S. Mukherjee, *Chem.–Eur. J.*, 2014, **20**, 12483.
- 115 S. Ganta and D. K. Chand, *Dalton Trans.*, 2015, **44**, 15181.
- 116 M. Han, Y. Luo, B. Damaschke, L. Gómez, X. Ribas, A. Jose, P. Peretzki, M. Seibt and G. H. Clever, *Angew. Chem., Int. Ed.*, 2015, **55**, 445.
- 117 S. Lim, H. Moon, D. Kim and O.-S. Jung, *Dalton Trans.*, 2021, **50**, 14320.
- 118 T. Zhang, L.-P. Zhou, X.-Q. Guo, L.-X. Cai and Q.-F. Sun, *Nat. Commun.*, 2017, **8**, ncomms15898.
- 119 J. Tessarolo, H. Lee, E. Sakuda, K. Umakoshi and G. H. Clever, *J. Am. Chem. Soc.*, 2021, **143**, 6339.
- 120 A. Kumar, R. Banerjee, E. Zangrando and P. S. Mukherjee, *Inorg. Chem.*, 2022, **61**, 2368.
- 121 K. Wu, E. Benchimol, A. Baksi and G. H. Clever, *Nat. Chem.*, 2024, **16**, 584.
- 122 M. D. Wise, J. J. Holstein, P. Pattison, C. Besnard, E. Solari, R. Scopelliti, G. Bricogne and K. Severin, *Chem. Sci.*, 2015, **6**, 1004.
- 123 S. M. Jansze, G. Cecot, M. D. Wise, K. O. Zhurov, T. K. Ronson, A. M. Castilla, A. Finelli, P. Pattison, E. Solari, R. Scopelliti, G. E. Zelinskii, A. V. Vologzhanina, Y. Z. Voloshin, J. R. Nitschke and K. Severin, *J. Am. Chem. Soc.*, 2016, **138**, 2046.
- 124 C. Klein, C. Gütz, M. Bogner, F. Topić, K. Rissanen and A. Lützen, *Angew. Chem., Int. Ed.*, 2014, **53**, 3739.
- 125 K. Suzuki, M. Tominaga, M. Kawano and M. Fujita, *Chem. Commun.*, 2009, 1638.
- 126 E. O. Bobylev, D. A. Poole, B. de Bruin and J. N. H. Reek, *J. Am. Chem. Soc.*, 2022, **144**, 15633.
- 127 C. Gütz, R. Hovorka, C. Klein, Q. Jiang, C. Bannwarth, M. Engeser, C. Schmuck, W. Assenmacher, W. Mader, F. Topić, K. Rissanen, S. Grimme and A. Lützen, *Angew. Chem., Int. Ed.*, 2014, **53**, 1693.
- 128 B. Roy, R. Saha, A. K. Ghosh, Y. Patil and P. S. Mukherjee, *Inorg. Chem.*, 2017, **56**, 3579.
- 129 M. Frank, M. D. Johnstone and G. H. Clever, *Chem.–Eur. J.*, 2016, **22**, 14104.
- 130 M. Fukuda, R. Sekiya and R. Kuroda, *Angew. Chem., Int. Ed.*, 2008, **47**, 706.
- 131 S. Freye, J. Hey, A. Torras-Galán, D. Stalke, R. Herbst-Irmer, M. John and G. H. Clever, *Angew. Chem., Int. Ed.*, 2012, **51**, 2191.
- 132 T. R. Schulte, M. Krick, C. I. Asche, S. Freye and G. H. Clever, *RSC Adv.*, 2014, **4**, 29724.
- 133 M. Frank, J. Hey, I. Balcioglu, Y. Chen, D. Stalke, T. Suenobu, S. Fukuzumi, H. Frauendorf and G. H. Clever, *Angew. Chem., Int. Ed.*, 2013, **52**, 10102.
- 134 T. R. Schulte, J. J. Holstein and G. H. Clever, *Angew. Chem., Int. Ed.*, 2019, **58**, 5562.
- 135 S. Freye, R. Michel, D. Stalke, M. Pawliczek, H. Frauendorf and G. H. Clever, *J. Am. Chem. Soc.*, 2013, **135**, 8476.
- 136 S. Watanabe, M. Yamashina, E. Tsurumaki and S. Toyota, *ChemistryEurope*, 2023, **1**, e202300047.
- 137 R. Zhu, J. Lübben, B. Dittrich and G. H. Clever, *Angew. Chem., Int. Ed.*, 2015, **54**, 2796.
- 138 J. I. Virtue, S. Tsoukatos, M. R. Johnston and W. M. Bloch, *Chem. Sci.*, 2024, **15**, 19119.
- 139 T. Abe, Y. Zhang, K. Takeuchi and S. Hiraoka, *Chem*, 2025, **11**, 102453.
- 140 Y. Yamauchi, M. Yoshizawa and M. Fujita, *J. Am. Chem. Soc.*, 2008, **130**, 5832.
- 141 T. K. Ronson, Y. Wang, K. Baldrige, J. S. Siegel and J. R. Nitschke, *J. Am. Chem. Soc.*, 2020, **142**, 10267.
- 142 P. Montes-Tolentino, A. S. Mikherdov, C. Drechsler, J. J. Holstein and G. H. Clever, *Angew. Chem., Int. Ed.*, 2025, **64**, e202423810.
- 143 D. M. Engelhard, S. Freye, K. Grohe, M. John and G. H. Clever, *Angew. Chem., Int. Ed.*, 2012, **51**, 4747.
- 144 S. Sudan, F. Fadaei-Tirani, R. Scopelliti, K. E. Ebbert, G. H. Clever and K. Severin, *Angew. Chem., Int. Ed.*, 2022, **61**, e202201823.
- 145 C. Hu, D. W. Chen, S. Sudan and K. Severin, *Chem. Sci.*, 2025, **16**, 5559.
- 146 M. Käseborn, J. J. Holstein, G. H. Clever and A. Lützen, *Angew. Chem., Int. Ed.*, 2018, **57**, 12171.
- 147 W. M. Bloch, J. J. Holstein, B. Dittrich, W. Hiller and G. H. Clever, *Angew. Chem., Int. Ed.*, 2018, **57**, 5534.



- 148 W. M. Bloch and G. H. Clever, *Chem. Commun.*, 2017, **53**, 8506.
- 149 M. Frank, L. Krause, R. Herbst-Irmer, D. Stalke and G. H. Clever, *Dalton Trans.*, 2014, **43**, 4587.
- 150 M. Yoshizawa, M. Nagao, K. Kumazawa and M. Fujita, *J. Organomet. Chem.*, 2005, **690**, 5383.
- 151 A. K. Bar, G. Mostafa and P. S. Mukherjee, *Inorg. Chem.*, 2010, **49**, 7647.
- 152 A. K. Bar, S. Raghothama, D. Moon and P. S. Mukherjee, *Chem.–Eur. J.*, 2012, **18**, 3199.
- 153 D. Samanta, S. Shanmugaraju, S. A. Joshi, Y. P. Patil, M. Nethaji and P. S. Mukherjee, *Chem. Commun.*, 2012, **48**, 2298.
- 154 D. Preston, J. E. Barnsley, K. C. Gordon and J. D. Crowley, *J. Am. Chem. Soc.*, 2016, **138**, 10578.
- 155 A. M. Johnson and R. J. Hooley, *Inorg. Chem.*, 2011, **50**, 4671.
- 156 R. Li, F. Fadaei-Tirani, R. Scopelliti and K. Severin, *Chem.–Eur. J.*, 2021, **27**, 9439.
- 157 B. Chen, J. J. Holstein, A. Platzek, L. Schneider, K. Wu and G. H. Clever, *Chem. Sci.*, 2022, **13**, 1829.
- 158 P. J. Stang and B. Olenyuk, *Acc. Chem. Res.*, 1997, **30**, 502.
- 159 J.-R. Li and H.-C. Zhou, *Nat. Chem.*, 2010, **2**, 893.
- 160 Q. Sun, S. Sato and M. Fujita, *Angew. Chem., Int. Ed.*, 2014, **53**, 13510.
- 161 S. Saha, B. Holzapfel, Y.-T. Chen, K. Terlinden, P. Lill, C. Gatsogiannis, H. Rehage and G. H. Clever, *J. Am. Chem. Soc.*, 2018, **140**, 17384.
- 162 K. Wu, J. Tessarolo, A. Baksi and G. H. Clever, *Angew. Chem., Int. Ed.*, 2022, **61**, e202205725.
- 163 J. Tessarolo, E. Benchimol, A. Jouaiti, M. W. Hosseini and G. H. Clever, *Chem. Commun.*, 2023, **59**, 3467.
- 164 K. Wu, B. Zhang, C. Drechsler, J. J. Holstein and G. H. Clever, *Angew. Chem., Int. Ed.*, 2021, **60**, 6403.
- 165 A. Platzek, S. Juber, C. Yurtseven, S. Hasegawa, L. Schneider, C. Drechsler, K. E. Ebbert, R. Rudolf, Q. Yan, J. J. Holstein, L. V. Schäfer and G. H. Clever, *Angew. Chem., Int. Ed.*, 2022, **61**, e202209305.
- 166 A. Tarzia, W. Shan, V. Posligua, C. J. T. Cox, L. Male, B. D. Egleston, R. L. Greenaway, K. E. Jelfs and J. E. M. Lewis, *Chem.–Eur. J.*, 2025, **31**, e202403336.
- 167 K. E. Ebbert, F. Sendzik, L. Neukirch, L. Eberlein, A. Platzek, P. Kibies, S. M. Kast and G. H. Clever, *Angew. Chem., Int. Ed.*, 2024, **64**, e202416076.
- 168 T. Abe, N. Sanada, K. Takeuchi, A. Okazawa and S. Hiraoka, *J. Am. Chem. Soc.*, 2023, **145**, 28061.
- 169 R.-J. Li, J. Tessarolo, H. Lee and G. H. Clever, *J. Am. Chem. Soc.*, 2021, **143**, 3865.
- 170 W.-T. Dai, T.-T. Liu, Q. Bai, Z. Zhang, P. Wang, W. Lu and Q. Zhang, *Sci. China Chem.*, 2024, **67**, 4110.
- 171 Z. Guo, H. Yu, J. Shi, N. Han, G. Wu, H. Zhang, B. Li and M. Wang, *Angew. Chem., Int. Ed.*, 2025, **64**, e202425369.
- 172 Y. Liu, S. Liao, W. Dai, Q. Bai, S. Lu, H. Wang, X. Li, Z. Zhang, P. Wang, W. Lu and Q. Zhang, *Angew. Chem., Int. Ed.*, 2023, **62**, e202217215.
- 173 E. Benchimol, I. Regeni, B. Zhang, M. Kabiri, J. J. Holstein and G. H. Clever, *J. Am. Chem. Soc.*, 2024, **146**, 6905.
- 174 S. Prusty, S. Krishnaswamy, S. Bandi, B. Chandrika, J. Luo, J. S. McIndoe, G. S. Hanan and D. K. Chand, *Chem.–Eur. J.*, 2015, **21**, 15174.
- 175 S. Prusty, K. Yazaki, M. Yoshizawa and D. K. Chand, *Chem.–Eur. J.*, 2017, **23**, 12456.
- 176 P. Howlader, P. Das, E. Zangrando and P. S. Mukherjee, *J. Am. Chem. Soc.*, 2016, **138**, 1668.
- 177 S. Sudan, R.-J. Li, S. M. Jansze, A. Platzek, R. Rudolf, G. H. Clever, F. Fadaei-Tirani, R. Scopelliti and K. Severin, *J. Am. Chem. Soc.*, 2021, **143**, 1773.
- 178 S. Samantray, S. Krishnaswamy and D. K. Chand, *Nat. Commun.*, 2020, **11**, 880.
- 179 P. Molinska, A. Tarzia, L. Male, K. E. Jelfs and J. E. M. Lewis, *Angew. Chem., Int. Ed.*, 2023, **62**, e202315451.
- 180 R. A. S. Vasdev, D. Preston, C. A. Casey-Stevens, V. Martí-Centelles, P. J. Lusby, A. L. Garden and J. D. Crowley, *Inorg. Chem.*, 2023, **62**, 1833.
- 181 S. K. Sen and R. Natarajan, *Inorg. Chem.*, 2019, **58**, 7180.
- 182 B. E. Barber, E. M. G. Jamieson, L. E. M. White and C. T. McTernan, *Chem*, 2024, **10**, 2792.
- 183 D. F. Brightwell, K. Samanta, J. Muldoon, P. C. Fleming, Y. Ortin, L. Mardiana, P. G. Waddell, M. J. Hall, E. R. Clark, F. Fantuzzi and A. Palma, *ChemistryEurope*, 2025, **3**, e202400050.
- 184 O. Jurček, P. Bonakdarzadeh, E. Kalenius, J. M. Linnanto, M. Groessl, R. Knochenmuss, J. A. Ihalainen and K. Rissanen, *Angew. Chem., Int. Ed.*, 2015, **54**, 15462.
- 185 O. Jurček, Nonappa, E. Kalenius, P. Jurček, J. M. Linnanto, R. Puttreddy, H. Valkenier, N. Houbenov, M. Babiak, M. Peterek, A. P. Davis, R. Marek and K. Rissanen, *Cell Rep. Phys. Sci.*, 2021, **2**, 100303.
- 186 O. Jurček, S. Chattopadhyay, E. Kalenius, J. M. Linnanto, A. Kiesilä, P. Jurček, P. Radiměřský and R. Marek, *Angew. Chem., Int. Ed.*, 2024, **63**, e202409134.
- 187 J. E. M. Lewis, A. Tarzia, A. J. P. White and K. E. Jelfs, *Chem. Sci.*, 2020, **11**, 677.
- 188 H. Yu, J. Li, C. Shan, T. Lu, X. Jiang, J. Shi, L. Wojtas, H. Zhang and M. Wang, *Angew. Chem., Int. Ed.*, 2021, **60**, 26523.
- 189 A. Tarzia, J. E. M. Lewis and K. E. Jelfs, *Angew. Chem., Int. Ed.*, 2021, **60**, 20879.
- 190 D. Ogata and J. Yuasa, *Angew. Chem., Int. Ed.*, 2019, **58**, 18424.
- 191 S. S. Mishra, S. V. K. Kompella, S. Krishnaswamy, S. Balasubramanian and D. K. Chand, *Inorg. Chem.*, 2020, **59**, 12884.
- 192 S. S. Mishra and D. K. Chand, *Dalton Trans.*, 2022, **51**, 11650.
- 193 J. E. M. Lewis, *Chem.–Eur. J.*, 2021, **27**, 4454.
- 194 M. Parbin, V. Sivalingam and D. K. Chand, *Angew. Chem., Int. Ed.*, 2024, **63**, e202410219.
- 195 R.-J. Li, A. Marcus, F. Fadaei-Tirani and K. Severin, *Chem. Commun.*, 2021, **57**, 10023.
- 196 J. de Montmollin, A. B. Solea, D. W. Chen, F. Fadaei-Tirani and K. Severin, *Inorg. Chem.*, 2024, **63**, 4583.
- 197 P. Molinska, L. Male and J. E. M. Lewis, *ChemRxiv*, 2024, DOI: [10.26434/chemrxiv-2024-0thnh](https://doi.org/10.26434/chemrxiv-2024-0thnh).



- 198 Z. T. Avery, M. G. Gardiner and D. Preston, *Angew. Chem., Int. Ed.*, 2025, **64**, e202418079.
- 199 L.-J. Chen, H.-B. Yang and M. Shionoya, *Chem. Soc. Rev.*, 2017, **46**, 2555.
- 200 L. M. Mesquita, J. Anhäuser, D. Bellaire, S. Becker, A. Lützen and S. Kubik, *Org. Lett.*, 2019, **21**, 6442.
- 201 X. Tang, H. Jiang, Y. Si, N. Rampal, W. Gong, C. Cheng, X. Kang, D. Fairen-Jimenez, Y. Cui and Y. Liu, *Chem*, 2021, **7**, 2771.
- 202 D. R. Martir, L. Delforce, D. B. Cordes, A. M. Z. Slawin, S. L. Warriner, D. Jacquemin and E. Zysman-Colman, *Inorg. Chem. Front.*, 2020, **7**, 232.
- 203 D. R. Martir, D. Escudero, D. Jacquemin, D. B. Cordes, A. M. Z. Slawin, H. A. Fruchtl, S. L. Warriner and E. Z. Colman, *Chem.–Eur. J.*, 2017, **23**, 14358.
- 204 C. Gütz, R. Hovorka, G. Schnakenburg and A. Lützen, *Chem.–Eur. J.*, 2013, **19**, 10890.
- 205 S. Ghorai, S. Maji, B. Paul, K. Samanta, S. K. Sen and R. Natarajan, *Chem.–Asian J.*, 2023, **18**, e202201312.
- 206 S. Ghorai and R. Natarajan, *Chem.–Eur. J.*, 2023, **29**, e202203085.
- 207 C. Ma, Q. Cao, L. Yu, Z. Ma and Q. Gan, *Angew. Chem., Int. Ed.*, 2024, **63**, e202410731.
- 208 R.-J. Li, J. de Montmollin, F. Fadaei-Tirani, R. Scopelliti and K. Severin, Construction of Pd-based coordination cages with three geometrically distinct ligands, *Dalton Trans.*, 2023, **52**, 6451–6456.

