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Introduction

Organosilicon chemistry plays a central role in organic synthesis and materials science because silicon is earth-abundant and displays peculiar properties, including a relatively low electronegativity and a high oxophilicity. Among the myriad of organosilicon compounds available for synthesis, those combining a functionalized organic moiety (“R”) bound to an alkoxy group (*e.g.*, R–SiMe_{3–n}(OR)_n, 1 ≤ n ≤ 3 with R = alkyl) are particularly attractive because they display an application potential (Scheme 1A) much greater than the corresponding triorganosilyl derivatives (*e.g.*, R–SiMe₃)¹ while being more stable than the corresponding halosilane derivatives (*e.g.*, R–SiCl₃). On the one hand, the alkoxy groups can be readily displaced in hydrolytic or non-hydrolytic sol/gel processes to yield functionalized polysiloxane (n = 2) or polysilsesquioxane (n = 3; R–SiO_{1.5}) materials. On the other hand, the electron-withdrawing alkoxy groups facilitate the formation of silicate intermediates that are key species for example in the Hiyama cross-coupling reaction,² the Tamao–Fleming oxidation³ or other oxidative processes.⁴

To date, the direct introduction of alkoxy silyl groups onto organic substrates has primarily been achieved through two main strategies: (1) the catalytic hydrosilylation of alkenes or alkynes with alkoxy silanes (*e.g.*, HSi(OEt)₃)⁵ and (2) the stoichiometric coupling of organolithium or Grignard reagents with chloro(alkoxy)silanes or orthosilicates (Si(OR)₄). Because

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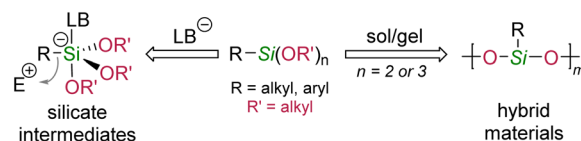
Catalytic alkoxy silylation of C–H bonds with *tert*-butyl-substituted alkoxy silyldiazenes†

Lamine Saadi, Loïc Valade and Clément Chauvier *

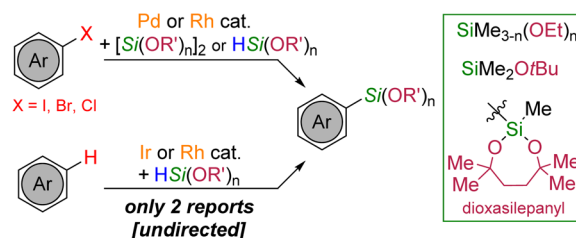
Organoalkoxy silanes (*e.g.* R–SiMe_{3–n}(OR)_n, 1 ≤ n ≤ 3 with R = alkyl or aryl) have found various applications in synthetic chemistry and materials science because the silicon-bound alkoxy groups provide unique opportunities for further derivatization and transformations. Among the few catalytic strategies that allow the direct and intermolecular introduction of an alkoxy silyl unit onto an organic substrate, the alkoxy silylation of unactivated C–H bonds has barely been achieved despite its synthetic potential and the atom-economy it conveys. In particular, a catalytic and transition metal-free C–H silylation protocol towards this class of organosilicon compounds has yet to be reported. We herein describe the first general alkoxy silylation of (hetero)arene C(sp²)–H and benzylic C(sp³)–H bonds under ambient, transition metal-free conditions using newly-prepared *tert*-butyl-substituted alkoxy silyldiazenes (tBu–N=N–SiMe_{3–n}(OR)_n, 1 ≤ n ≤ 3 with R = Et, *i*Pr or *t*Bu) as silylating reagents and *t*BuOK as catalytic promoter.

the hydrosilylation approach is effective only for synthesizing alkyl- and vinylsilane derivatives, the stoichiometric route remains indispensable for preparing aryl- and primary benzyloxy silanes.⁶ However, the latter methods suffer from several drawbacks, including low tolerance for electrophilic functional

A. Applications of organo(alkoxy)silanes

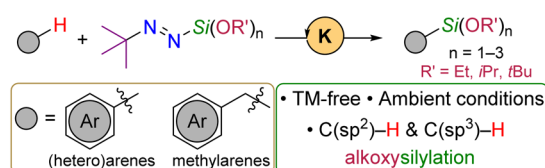


B. TM-catalyzed C(sp²)–Si(OR)_n cross-couplings



• Noble metals • high temperature • low generality

C. TM-free catalytic C–Si(OR)_n cross-couplings



Scheme 1 Applications and catalytic synthesis of organo(alkoxy) silanes.



groups and the need for cryogenic temperatures to control reactivity and selectivity. Furthermore, they generate substantial amounts of metallic waste, posing significant challenges for large-scale applications.

In recent years, transition metal-catalyzed C–Si cross-couplings have been reported to introduce mono-,⁷ bis-⁸ and trisalkoxy⁹ as well as siloxy¹⁰ units onto aromatic scaffolds directly from aryl halides and the corresponding hydrosilane or disilane derivatives (Scheme 1B). However, not only are those approaches based on expensive noble metals (Pd or Rh), but they also rely on pre-functionalized halide substrates, selective preparation of which can be difficult and reduces the overall atom economy of the silylation process.

For these reasons, the direct intermolecular silylation of C–H bonds has emerged as a privileged strategy to forge C(sp²)– or C(sp³)–Si bonds in a catalytic manner.¹¹ While this field has substantially evolved over the past decade with useful solutions proposed to overcome both reactivity and site-selectivity issues,¹² the general introduction of alkoxy-silyl units rather than trialkyl or siloxy-based (SiMe(OSiMe₃)₂) ones onto hydrocarbon substrates has barely been achieved. As far as arylsilanes are concerned, only two such reports have been described with hydrosilanes as silicon sources (Scheme 1B). Lee and coworkers showed that a rhodium-based complex catalyzes the undirected silylation of unactivated (hetero)arenes with HSiMe₂OEt as the alkoxy-silyl source.¹³ This catalytic system, however, proved inefficient to promote the transfer of alkoxy-silyl groups bearing more than one ethoxy unit and the silylated arenes are mostly produced as mixtures of regioisomers. In 2021, Yorimitsu and Shimokawa showed that Hartwig's catalytic dehydrogenative silylation protocol could be modified to introduce the dioxasilepanyl unit, a hydrolytically-stable 7-membered cyclic bisalkoxy-silyl group.¹⁴ Nevertheless, this iridium-catalysed transformation, initially cantoned to the exclusive transfer of the bis-siloxy SiMe(OSiMe₃)₂ unit,¹⁵ could not be further extended to acyclic alkoxy-silyl groups derived from common alcohols.

Beyond the alkoxy-silylated arenes, benzylsilanes bearing alkoxy-silyl units are also promising monomers in silicone science,¹⁶ which have barely been explored because of the lack of general synthesis. In particular, the catalytic and undirected formation of benzyl(alkoxy)silanes by dehydrogenative silylation of toluene derivatives with HSi(OEt)₃ has only been reported to occur with heterogeneous metal-based catalysts at high temperature.¹⁷

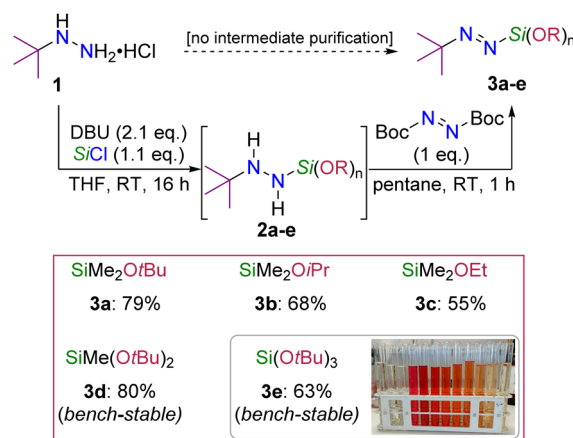
From these considerations, we have sought to develop a general method towards the formation of aryl- and benzyl-(alkoxy)silanes through the silylation of unactivated (hetero) arene C(sp²)–H bonds and benzylic C(sp³)–H bonds. To that end, we have recently reported that potent metalating agents can be generated catalytically when *tert*-butyl-substituted silyl diazenes (*t*Bu–N=N–SiR₃) are treated with a catalytic amount of a potassium salt (typically *t*BuOK). This process has been fruitfully exploited to introduce trialkylsilyl (*e.g.*, SiMe₃ and SiEt₃) groups through the catalytic silylation of unactivated C(sp²)–H¹⁸ and C(sp³)–H bonds¹⁹ *via* metalation/silylation sequences. We herein demonstrate that such a reactivity

manifold can now be applied to the catalytic introduction of alkoxy-silyl units from newly-prepared *tert*-butyl-substituted alkoxy-silyl diazenes under ambient and transition metal-free conditions (Scheme 1C).

Results and discussion

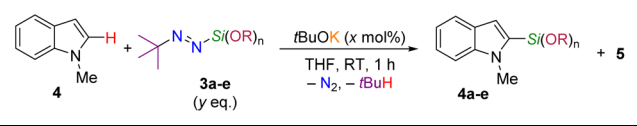
We initiated our study with the preparation of various *tert*-butyl-substituted alkoxy-silyldiazenes (*t*Bu–N=N–Si(OR)_{*n*}) by varying the nature (R = Et, *i*Pr and *t*Bu) and/or the number (*n* = 1, 2 or 3) of alkoxy units bound to the silicon atom. To that end, our previously reported protocol starting from commercially-available *tert*-butylhydrazine hydrochloride **1** could largely be used with only marginal modification of the conditions needed in some cases (see the ESI† for details), further demonstrating the robustness of this process.^{18,20} Typically, **1** was mixed with a chloro(alkoxy)silane Cl–SiMe_{3–*n*}(OR)_{*n*} (*n* = 1, 2 or 3) under basic conditions to afford the corresponding (alkoxy)silylated hydrazines **2a–e** arising from the selective replacement of the chloride by the hydrazide moiety (Scheme 2). The subsequent dehydrogenative oxidation of the non-purified crude hydrazine with DBAD (di-*tert*-butyl azodicarboxylate)²¹ produced the desired red-colored alkoxy-silyldiazenes **3a–e** bearing one (**3a–3c**), two (**3d**) or three (**3e**) alkoxy units in 55–80% yields over two steps. Importantly, the bulkiest bis- and tris-*tert*-butoxy derivatives **3d** and **3e** were found bench-stable and could be easily purified by column chromatography on deactivated (NEt₃) silica gel. In all instances, the *tert*-butyl-substituted alkoxy-silyldiazenes **3a–e** were found stable towards the intermolecular redistribution of the alkoxy/diazenyl groups (*e.g.* 2 *t*Bu–N=N–SiMe₂(OR) → (*t*Bu–N=N)₂–SiMe₂ + SiMe₂(OR)₂) upon storage, though the least bulky congener **3c** slowly decomposed over a period of months when kept at room temperature in a glovebox, presumably through such a redistribution pathway.

With these new silyldiazenes in hands, we first studied the transfer of the mono-alkoxy-silyl units SiMe₂OR from diazenes **3a–c** to *N*-methylindole (**4**) through the catalytic deprotonative silylation of the weakly acidic C(sp²)–H bond at the 2-position

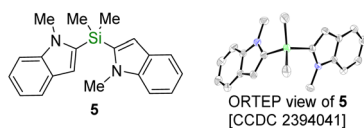


Scheme 2 Synthesis of *tert*-butyl-substituted alkoxy-silyldiazenes. Yields refer to those of isolated products over two steps (1 → 3a–e).



Table 1 Optimization of the C(sp³)-H silylation of **4** with diazene **3a-e**^a


Entry	x	Si(OR) _n [y]	ρ [%]	Yield 4a-e [%]	Yield 5 [%]
1	10	3a SiMe ₂ O <i>t</i> Bu [1.8]	>95	88 (91)	7
2	10	3b SiMe ₂ O <i>i</i> Pr [1.8]	70	51	9
3	10	3c SiMe ₂ OEt [1.8]	65	42	10
4	10	3b SiMe ₂ O <i>i</i> Pr [3.0]	86	73 (49)	7
5	10	3c SiMe ₂ OEt [3.0]	69	63	4
6	50	3a SiMe ₂ O <i>t</i> Bu [2.5]	91	73	3
7	50	3b SiMe ₂ O <i>i</i> Pr [2.5]	73	7	40
8	50	3c SiMe ₂ OEt [2.5]	73	3	50 (48)
9 ^b	10	3d SiMe(O <i>t</i> Bu) ₂ [2.2]	—	(99)	n.a.
10 ^c	40	3e Si(O <i>t</i> Bu) ₃ [2.2]	75	68 (46)	n.a.



^a Conversion (ρ) and yields were determined by ¹H NMR on the crude reaction mixture with 1,3,5-trimethoxybenzene as internal standard. Isolated yields in brackets. n.a.: not applicable. ^b 16 h. ^c 24 h.

(Table 1, entries 1–8). To that end, the reaction parameters were first optimized with the bulkiest congener **3a** (see Table S1 in the ESI† for details). Using *t*BuOK (10 mol%) as an inexpensive catalytic promoter and a slight excess of **3a** (1.8 eq.) in THF, **4** was rapidly and cleanly converted to the corresponding silylated indole **4a**, which was isolated in 91% yield after 1 h of reaction at room temperature (entry 1, Table 1). Careful analysis by NMR spectroscopy and GC-MS of the crude reaction mixture prior to purification revealed that the remaining mass balance was mainly accounted for by the presence of the chromatography-separable bis-indole product **5** (ca. 7%), formulation of which was eventually confirmed by single-crystal X-ray diffraction (see bottom of Table 1). **5** formally arises from a substitution of the *tert*-butoxy group in **4a** by a second 2-indolyl moiety.

Under otherwise identical conditions, the less bulky isopropoxy and ethoxy diazenes **3b** and **3c** reacted more sluggishly as the conversion plateaued at 70% and 65% and the corresponding indole products **4b** and **4c** were produced in 51% and 42% yield along with 9% and 10% of **5**, respectively (entries 2–3). In both cases, the diazenes **3b** and **3c** were fully consumed prior to reaching full conversion of **4**, thereby indicating that a non-productive decomposition occurred in the presence of *t*BuOK. Accordingly, in the absence of **4**, **3b** and **3c** rapidly reacted (<5 min) with *t*BuOK (10 mol%) to afford a range of silylated products, among which silylated di-*tert*-butylhydrazines were produced predominantly (see Fig. S5 in the ESI†). In line with these findings, higher yields of **4b** and **4c** (73% and 63%) were obtained when **4** was treated with increased amounts

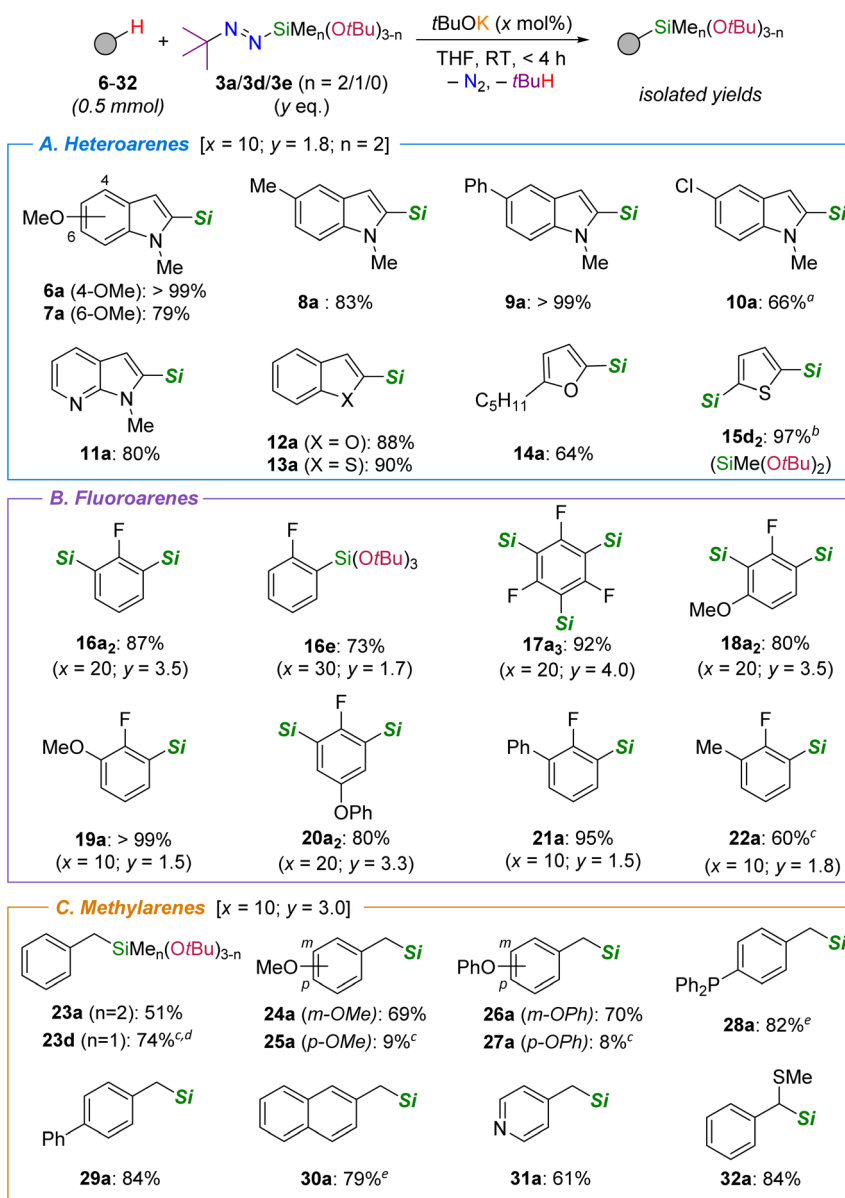
of **3b** and **3c** (3 eq.) to compensate for their non-productive self-degradation (entries 4–5). Taken together, these results indicate that the C(sp²)-H silylation reaction benefits from an increased steric bulk of the alkoxide residue, primarily because of the greater stability of the corresponding diazene towards its self-decomposition. In addition, the formation of the doubly-substituted product **5** remains marginal with diazene **3a**, while it becomes the major product at higher loading of *t*BuOK with **3b** and **3c** (entries 6–8).

In contrast to the mono-alkoxysilyldiazenes **3a-c**, the bench-stable bis- and tris-*tert*-butoxysilyl derivatives **3d** and **3e** were found highly resistant to base-catalyzed self-decomposition and barely prone to over-substitution. As an illustration, the product **4e** was isolated in quantitative yield after 16 h of reaction and even the extremely bulky Si(O*t*Bu)₃ unit could be transferred in 68% yield after ca. 24 h of reaction (75% conversion) with 40 mol% *t*BuOK (entries 9–10). Importantly, the latter result constitutes the first example of a synthesis of an aromatic tri-alkoxysilane by catalytic silylation of a C(sp²)-H bond.

We next turned our attention to the substrate scope of the silylation reaction (Scheme 3), mostly using diazene **3a** as the silicon source not only because it performs better than its less bulky congeners **3b** and **3c**, but also because the SiMe₂O*t*Bu unit is stable towards protonolysis on silica gel, thereby simplifying product purification and isolation. At the outset, various indoles bearing electron-donating (6–8) or withdrawing groups (9–10) on the arene core could be silylated in good to excellent yields under the conditions optimized with **4** (10 mol% *t*BuOK, 1.8 eq. **3a**). In all instances the silylation occurred selectively at the 2-position even when potentially directing methoxy (in 6–7) and chloride (in 10) groups that acidify proximate hydrogen atoms are present. Likewise, the silylation of the azaindole derivative **11** cleanly delivered **11a** in 80% isolated yield. Beyond the indole core, other heteroarenes were also efficiently silylated, including furan (**12** & **14**) or thiophene (**13** & **15**) derivatives. Importantly, the bis-silylation of thiophene **15** could easily be achieved with diazene **3d** (3.2 eq.) to give **15d₂** in excellent yield, a bifunctional organosilane that may find applications as monomer in silicone science.

Simple fluoroarenes were also competent substrates (Scheme 3B). As an illustration, one of the *ortho* C-H bonds in fluorobenzene **16**, which displays a similar acidity than the one at the 2-position of *N*-methylindole (*pK*_a ≈ 37 in DMSO),²² was readily silylated with *t*BuOK (10 mol%) and **3a** (1.2 eq.) to afford the corresponding SiMe₂O*t*Bu-substituted arene **16a** in 67% yield (not shown). However, under these conditions, the latter was produced along with 24% of the bis-silylated arene **16a₂** as judged by ¹⁹F NMR spectroscopy, which could be isolated in 87% yield when an excess of **3a** (3.5 eq.) was employed. In fact, we had previously shown that the selectivity for mono- vs. bis-silylation pathways with trialkylsilyl-substituted diazenes can be controlled by tuning the steric hindrance of the silyl group (e.g. Me₃Si vs. *t*BuMe₂Si).¹⁸ A similar dichotomy was observed with alkoxysilyl-substituted diazenes (see Table S6 in the ESI†). For example, using diazene **3e** as the silicon source, only one bulky Si(O*t*Bu)₃ unit could be selectively introduced in product **16e**, which was further isolated in 73% yield. In fact, the present





Scheme 3 Substrate scope of the silylation reaction ($Si = \text{SiMe}_2\text{OtBu}$ from diazene **3a**, unless otherwise indicated). Yields refer to those of isolated product after purification by flash chromatography.^a **3a** (2 eq.), $t\text{BuOK}$ (20 mol%).^b **3d** (3.2 eq.).^c Yield determined by NMR spectroscopy.^d **3d** (3 eq.).^e $t\text{BuOK}$ (20 mol%).

silylation methodology is particularly efficient to assemble polysilylated compounds as witnessed by the per-silylation of 1,3,5-trifluorobenzene (**17**), which was readily converted into **17a₃** (92% yield) with a slight excess of diazene **3a** (4 eq.). Interestingly, this type of trifunctional arylsilane product has been employed as precursor of siloxane-based cyclophanes.²³ The superb site selectivity for positions *ortho* to fluorine atom(s) was also retained for substrates bearing potentially competing methoxy (**18–19**) or phenoxy (**20**) directing groups as well as aryl (**21**) or methyl (**22**) substituents. In all instances, the corresponding mono- or bis-silylated products were isolated in high yields (>80%), though **22a** was accompanied by 16% of the bis-silylated product **22a₂**, in which both the $\text{C}(\text{sp}^2)\text{-H}$ bond *ortho* to fluorine and one benzylic $\text{C}(\text{sp}^3)\text{-H}$ bond have been silylated.

Capitalizing on this observation as well as on our previous work that dealt with the undirected silylation of benzylic $\text{C}(\text{sp}^3)\text{-H}$ bonds,^{19b} the lateral silylation of various methylarenes was investigated (Scheme 3C). When toluene (**23**) was reacted with diazene **3a** (3.0 eq.) in the presence of $t\text{BuOK}$ (10 mol%), the corresponding benzylsilane **23a** could be isolated in 51% yield after 4 h of reaction. With the bulkier diazene **3d**, however, the $\text{SiMe}(\text{OtBu})_2$ unit could be transferred to **23** in an improved 74% yield, thereby suggesting that increased steric hindrance (and Lewis acidity of the silicon center) minimizes pathways that shut down the catalytic activity. Nonetheless, other methylarenes bearing methoxy (**24**), phenoxy (**26**), diphenylphosphino (**28**) or phenyl (**29**) substituents reacted smoothly with **3a** to afford the corresponding silylated products in good to excellent



yields. In the cases of **24a** and **26a**, regioisomers arising from the silylation *ortho* to the MeO or PhO groups were barely observed, thus confirming that the high lateral selectivity we had previously demonstrated with Et₃Si-substituted diazene (*t*Bu-N=N-SiEt₃) also holds with **3a**. However, we have found one major reactivity difference between **3a** and *t*Bu-N=N-SiEt₃ when substrates bearing (strongly) electron-donating groups at the para position are employed. As an illustration, while the silylation of 4-methylanisole **25** proceeds smoothly with *t*Bu-N=N-SiEt₃ (84% isolated yield), only a small amount of **25a** was formed with diazene **3a** (<10% yield) under otherwise identical or more forcing conditions. A similar outcome was obtained with methylenes bearing 4-Oph (**27**) or 4-NMe₂ (not shown) substituents. Because meta-substituted isomers **24** and **26** are competent substrates, we ascribe the lack of reactivity of electron rich substrates to the relatively low acidity of the benzylic C(sp³)-H bonds, deprotonation of which would be potentially slower than side reactions, especially diazene self-decomposition. Despite these limitations, less aromatic scaffolds such as the naphthyl or pyridyl derivatives **30** and **31** as well as the secondary benzylic thioether **32** were also tolerated.

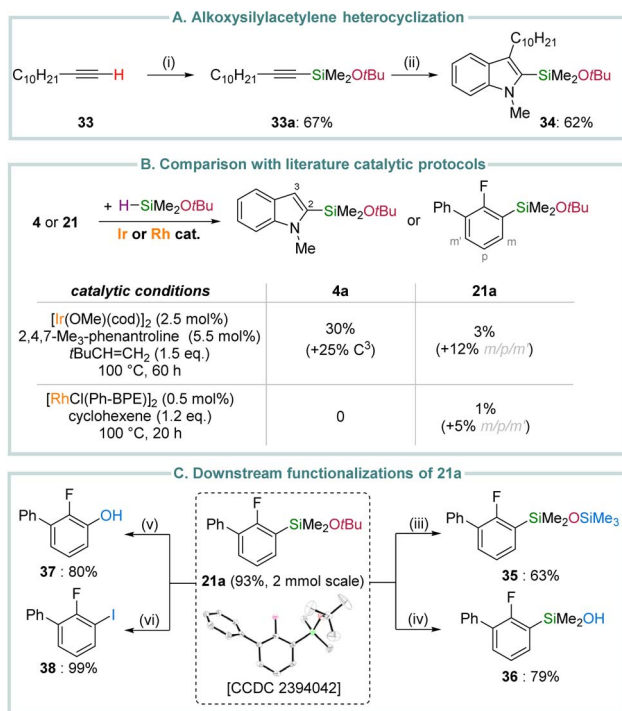
Beyond C(sp²)-H or C(sp³)-H bonds, our alkoxylation protocol can also be applied to the silylation of more acidic C(sp)-H bonds.²⁴ As an illustration, 1-dodecyne (**33**) smoothly reacted at room temperature with diazene **3a** (1.5 eq.) upon simple treatment with KOH (10 mol%) rather than *t*BuOK,

affording the silylated acetylene **33a** in 67% isolated yield (Scheme 4A). Interestingly, **33a** can be subjected to a Pd-catalysed Larock heterocyclization with 2-iodo-*N*-methylaniline according to a protocol developed by Denmark and Baird.²⁵ Such a route gives access to the 3-substituted 2-alkoxysilylated indole product **34** (62% yield), which can be further derivatized by Hiyama-Denmark cross-couplings.²⁵

Overall, the present catalytic methodology exhibits unique generality to assemble a range of structurally-diverse organosilanes bearing different alkoxy-silyl groups from readily available hydrocarbon substrates. In particular, it provides a direct entry to (hetero)aryl(*tert*-butoxy)silanes, a class of compounds that remains largely inaccessible through state-of-the-art C(sp²)-H alkoxy-silylation protocols. To illustrate this further, we reacted the model substrates **4** and **21** with HSiMe₂O*t*Bu under Rh-¹³ or Ir-based^{14a} catalytic conditions (Scheme 4B and Section 3 in the ESI† for details), which have been previously developed to introduce alkoxy-silyl units by (undirected) C(sp²)-H bond dehydrogenative silylation. Under such conditions, the desired products **4a** and **21a** were either not formed (Rh) or produced in low yields (Ir), despite the rather forcing conditions (100 °C), long reaction times (20–60 h) and the presence of hydrogen acceptors as additives. In addition, the Ir-based system displayed only moderate site-selectivity, as both **4a** and **21a** were accompanied by other regioisomers as well as bis-silylated products, as indicated by GC/MS analysis. In stark contrast, the potassium-catalysed silylation of **4** and **21** with **3a** proceeds efficiently at room temperature with superb site-selectivity.

The limited availability of catalytic methods to access (hetero)aryl(*tert*-butoxy)silanes has resulted in a scarcity of data regarding their potential applications.^{7b} This is particularly noteworthy given their bench stability and thus ease of handling compared to the more prevalent methoxy- or ethoxy-silyl derivatives.²⁶ We thus embarked into an evaluation of the application potential of the fluorobiphenyl product **21a** (Scheme 4C), which was prepared on a 2 mmol scale in 93% isolated yield (564 mg), further demonstrating the scalability of the methodology. Adopting a silicon chemistry point of view, we first demonstrated that despite its bulkiness, the *t*BuO residue can readily be replaced by other functional groups. For example, **21a** smoothly underwent a bismuth-catalyzed cleavage of the O-C(CH₃)₃ in the presence of chlorotrimethylsilane to afford the heteroleptic disiloxane **35** in 63% yield.²⁷ Alternatively, controlled hydrolysis of the O-Si bond in **21a** led to the silanol **36**.²⁵ Finally, oxidative cleavage of the C(sp²)-Si bond could be achieved to assemble the functionalized phenol **37** and iodoarene **38** in good to excellent yields.

To end this study, we sought to establish that the potassium-catalyzed C(sp²)-H and C(sp³)-H (alkoxy)silylation reactions proceed according to anionic mechanistic scenarios similar to those previously established with *tert*-butyl-substituted tri-alkylsilyldiazene.^{18,19,28} To that end, we first verified that under the typical C(sp²)-H silylation conditions, the *O*-fluoroaryl carbamate **39** undergoes an anionic *ortho*-Fries (AoF) rearrangement diagnostic of the involvement of aryl potassium intermediates. When **39** was reacted with **3a** (1.5 eq.) and



Scheme 4 Applications and interest of the methodology. (i): KOH (10 mol%), **3a** (1.5 eq.), RT, 3 h. (ii): Pd(OAc)₂ (5 mol%), PPh₃ (5 mol%), 2-iodo-*N*-methylaniline (1 eq.), **37a** (2 eq.), LiCl (1 eq.), K₂CO₃ (5 eq.), DMF, 100 °C, 5 h. (iii): BiCl₃ (5 mol%), Me₃SiCl (3 eq.), MeCN : pentane (3 : 1), RT, 3 h. (iv): HCl_(aq) (4 mol%), MeCN, RT, 10 min. (v): *t*BuOOH (2 eq.), *t*BuOK (2.2 eq.), THF, 60 °C, 16 h. (vi): ICl (3 eq.), CH₂Cl₂, RT, 16 h.



a stoichiometric amount of *t*BuOK (100 mol%), the rearranged potassium phenoxide product **40** was formed in almost quantitative yield (Scheme 5A). Under catalytic conditions, such a rearrangement was also operative as witnessed by the formation of the corresponding *O*-silylated benzamide (*ca.* 7% yield after 16 h, see Table S11 in the ESI†), which presumably arises from the silylation of the hydroxide of **40** in the presence of **3a**. Importantly, no reaction took place between **39** and *t*BuOK (1 eq.) in the absence of **3a**, thereby indicating that the diazene decomposition is pivotal to induce the metalation and the subsequent rearrangement of **39**. In a similar vein, the alkoxy-silylation of benzylic C(sp³)-H bonds follows an anionic mechanism as confirmed by the direct observation of the α -silylbenzylpotassium species **41** that acts as a resting state in the *t*BuOK-promoted reaction between toluene (**23**) and **3a** (Scheme 5B).

From this perspective, the (alkoxy)silylation of both C(sp²)-H and C(sp³)-H bonds proceeds through anionic chain reactions involving organopotassium intermediates as chain carriers (**I** and **II** in Scheme 5C). This reactivity manifold, comprising metalation (**I** → **II**) and silylation (**II** → **III**) events, does not radically differ from the one observed with trialkylsilyl-based diazenes, thereby demonstrating that the diazenyl unit (*t*Bu-N=N⁻) is a better leaving group than an alkoxide (RO⁻ with R = Et, *i*Pr or *t*Bu). In some instances, however, products arising from the nucleophilic displacement of both units have been detected (*e.g.*, the bis-indole **5**, see Table 1), especially when the alkoxide residue is sterically unhindered (R = Et or *i*Pr).

Although the exact pathway underpinning the formation of doubly-substituted products remains to be studied, the substitution of the alkoxy group in **III** by **II** may become kinetically accessible when the concentration of **III** is greater than that of the corresponding diazene (see Scheme 5C, dashed arrows). In line with this hypothesis, **5** was formed in 62% yield when the silylated indole **4b** was introduced from the outset of a reaction between indole **4** and diazene **3b** (Scheme 5D).

Conclusions

To conclude, we have demonstrated that the preparation of various organo(alkoxy)silanes can be accomplished through the potassium-catalyzed alkoxy-silylation of both aryl C(sp²)-H and benzylic C(sp³)-H bonds with newly-synthesized *tert*-butyl-substituted alkoxy-silyldiazenes. These operationally-simple transformations proceed under ambient, transition metal-free conditions and involve organopotassium compounds as reactive intermediates, high nucleophilicity of which allows the introduction of silyl groups as bulky as the *tris-tert*-butoxysilyl (Si(*Ot*Bu)₃). Because organoalkoxysilanes are widely used in synthetic chemistry and in materials science, it is our hope that the present methodology will stimulate further applications, notably owing to the easy access to bench-stable organosilanes bearing mono-, bis- or *tris-tert*-butoxysilyl units from readily available hydrocarbon substrates.

Data availability

The data supporting this article have been included as part of the ESI.† Crystallographic data for compounds **5** and **21a** have been deposited at the Cambridge Crystallographic Data Centre under accession numbers 2394041 and 2394042 and can be obtained free of charge from <http://www.ccdc.cam.ac.uk/structures>.

Author contributions

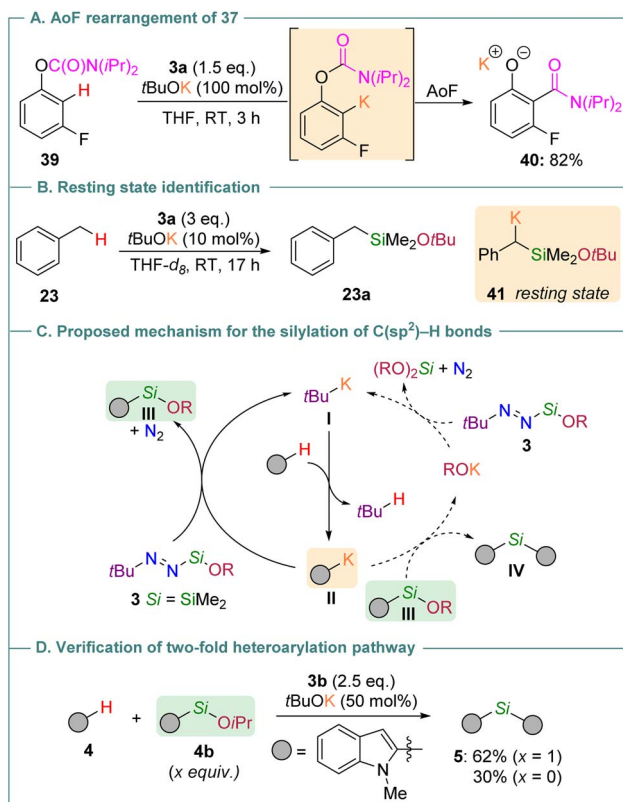
L. S. conducted all the experiments and analyzed the data. L. V. performed preliminary experiments. C. C. designed and directed the research and wrote the original draft. C. C. and L. S. wrote the ESI.†

Conflicts of interest

There are no conflicts to declare.

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Scheme 5 Mechanistic investigations.



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