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Leveraging heterocycle-fused 1,4-benzoquinone to design chemical modulators for both metal-free and metal-bound amyloid- β †

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The complex pathology of Alzheimer's disease includes various pathogenic components, such as metalfree amyloid-β (Aβ) and metal-bound Aβ (metal-Aβ). Here we report an effective strategy for developing novel heterocycle-fused 1,4-benzoquinone (BQ) compounds to control the aggregation and toxicity of both metal-free Aß and metal-Aß. We designed and synthesized these compounds by fusing BQ with 3pyrazolone responsible for metal chelation. The compounds' ability to form covalent bonds with $A\beta$ is tuned by the annulation of the BQ moiety and the type, position, and number of substituents on the 3pyrazolone group. Furthermore, the BQ functionality on the 3-pyrazolone framework can undergo ohydroxylation, enhancing its metal chelation in a bidentate manner. Our results demonstrate that these heterocycle-fused BQ compounds can redirect the assembly of AB into less toxic aggregates by binding to metal ions, modifying Aβ structures in both the absence and presence of metal ions, and promoting oxidative changes to Aß. This study highlights the importance of structural modifications and optimizations of BQ to leverage its strength of covalently cross-linking to AB and overcome its limitations in metal chelation and cytotoxicity, which are critical for designing chemical modulators for metal-free Aβ and metal-Aβ. Our approach offers a novel strategy for developing chemical modulators towards metal-related peptides and proteins as well as therapeutic agents for metal-associated amyloid disorders.

Introduction

Aberrant aggregation of peptides and proteins is linked to numerous pathological conditions.¹ The amyloid cascade hypothesis, which suggests that the accumulation of amyloid- β (A β) aggregates is a primary cause of Alzheimer's disease (AD), has been extensively studied. This has led to the development of various chemical tools and therapeutic agents aimed at targeting and modulating A β aggregation.² Notably, A β -directed monoclonal antibodies have been approved by the United States Food and Drug Administration as AD treatment, although concerns about their efficacy and safety persist. This underscores the importance of the amyloid paradigm in understanding the pathology of AD while highlighting the need for further investigation into other biological factors influencing A β aggregation profiles. High concentrations of metal ions, such as Cu(1/II) and Zn(II), are found in A β aggregates in the

As a means of modifying the aggregation profiles of A β , the manipulation of amino acid residues in the peptides has been reported. ^{22–25} As shown in Fig. 1a, 1,4-benzoquinone (BQ), capable of forming a covalent adduct with A β , has been demonstrated to alter the assembly routes of A β . These reactions, however, are initiated after A β accommodates certain small molecules [*e.g.*, *N,N*-dimethyl-*p*-phenylenediamine (DMPD), benzene-1,4-diamine, N^1 -((1*H*-pyrrol-2-yl)methyl)- N^4 , N^4 -dimethylbenzene-1,4-diamine (7), and N^1 , N^1 -dimethyl- N^4 -(thiophen-2-ylmethyl)benzene-1,4-diamine (L1)], which are then chemically transformed into BQ. ^{25,29-31} Moreover, BQ is unable to chelate metal ions, limiting its interaction with metal ions in metal–A β complexes. The inherent toxicity of BQ further restricts its biological applications. ^{25,32}

In this study, we illustrate an effective approach for developing novel heterocycle-fused **BQ** compounds that can control the aggregation and toxicity of both metal-free $A\beta$ and metal- $A\beta$. We rationally designed and synthesized heterocycle-fused **BQ** molecules and probed their effects on the assembly and

brains of AD patients.⁸⁻¹⁶ Metal coordination to A β peptides affects their conformation and aggregation pathways, indicating metal-bound A β (metal-A β) as an additional pathological component of AD.⁹⁻²¹ Therefore, developing chemical reagents that can modulate both metal-free A β and metal-A β is crucial for addressing the complex pathology of AD.

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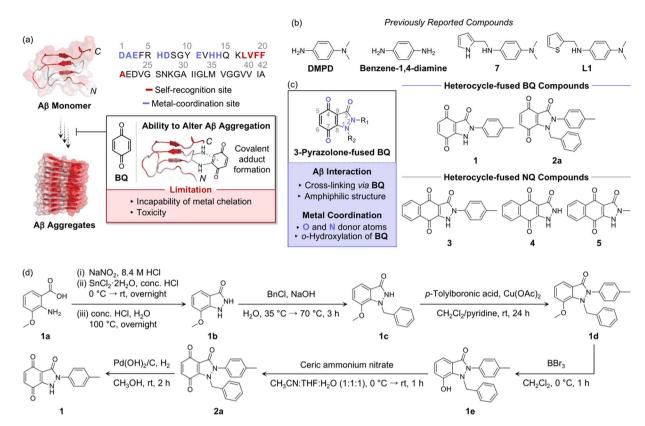


Fig. 1 Rational design and preparation of heterocycle-fused BQ compounds to modify the aggregation and toxicity of Aβ in the absence and presence of metal ions. (a) Ability of BQ to interact with A β and its properties that limit the reactivity of metal-A β and biological applications. Structures of monomeric (xx98156 26) and fibrillar (PDB 2BEG 27) $A\beta_{42}$ are shown according to the Eisenberg hydrophobicity scale (hydrophilic to hydrophobic amino acid residues colored in a gradient from white to red).28 The amino acid residues involved in the self-recognition site and metal coordination are highlighted in red and purple, respectively. (b) Previously reported compounds that can be transformed into BQ. (c) Chemical structures and properties of heterocycle-fused BQ and NQ compounds. 1, 2-(p-Tolyl)-1H-indazole-3,4,7(2H)-trione; 2a, 1-benzyl-2-(p-tolyl)-1H-indazole-3,4,7(2H)-trione; 3, 2-(p-tolyl)-1H-benzo[f]indazole-3,4,9(2H)-trione; 4, 1H-benzo[f]indazole-3,4,9(2H)-trione; 5, 2methyl-1H-benzo[f]indazole-3,4,9(2H)-trione. (d) Synthetic routes to 1 and 2a.

toxicity of metal-free A\beta and metal-A\beta. Additionally, we determined the molecular-level mechanisms for these reactivities, including their ability to form covalent cross-links with AB and chelate metal ions. Our study underscores the significance of structural alterations on the BQ moiety for creating chemical reagents and therapeutic agents with regulatory reactivities towards metal-related peptides and proteins associated with amyloid diseases.

Results and discussion

Design principle and synthesis of compounds

For the covalent adduct formation between AB and the BQ moiety, the peptide first needs to recruit BQ-containing compounds. To enhance direct interactions with both metalfree Aβ and metal-Aβ, 3-pyrazolone was fused into the BQ moiety, as illustrated in Fig. 1b. Pyrazolones are a class of pyrazole heterocycles with various chemical and biological properties.33-38 Particularly, the planarity of small molecules containing 3-pyrazolone combined with a benzene ring is suggested to favor the disruption of the β -sheet conformation of A β aggregates through hydrophobic interactions.39 The carbonyl

group on 3-pyrazolone also participates in hydrogen bonding with Aβ.39 Furthermore, the nitrogen (N) and oxygen (O) donor atoms from the amino and carbonyl groups on 3-pyrazolone can bind to metal ions. Upon fusion with BQ, there are two possible metal-chelation sites: (i) a pair of N and O donor atoms from the amino and carbonyl groups at 1- and 7-positions, respectively; (ii) two O donor atoms from the carbonyl groups at 3- and 4positions. Unlike the relatively weak metal binding of BQ through its carbonyl O donor atoms, metal-chelation properties of 3-pyrazolone-fused BQ enable a sufficient interaction with metal ions surrounded by Aß.

Through further functionalization of the 3-pyrazolone-fused BQ framework, we rationally designed heterocycle-fused BQ compounds (1 and 2a; Fig. 1b). The incorporation of a p-tolyl group at the 2-position, without interfering with the metalchelating moieties, generated compound 1 with increased hydrophobicity, which is suitable for hydrophobic interactions with Aβ. Previous studies have shown that a phenyl ring linked to the pyrazole moiety could engage in π - π , π -alkyl, and π cation interactions with the Phe, Val and Leu, and Lys residues, respectively, located near and in the self-recognition site (Fig. 1a) crucial in the early stages of Aβ aggregation. 40,41

Compound **2a**, with an additional benzyl group at the 1-position, was devised to further enhance the hydrophobicity of compound **1**, thereby improving hydrophobic contacts on A β and effectively controlling amyloid fibrillization. Although the benzyl group hinders the metal binding of the N donor atom at the 1-position, compound **2a** can still coordinate metal ions *via* O donor atoms from the carbonyl groups at the 3-, 4-, and 7-positions.

Since α, β -unsaturated carbonyl moieties in **BQ** are sensitive to nucleophilic attack at the β -carbon (C_{β}) atom, a Michael addition reaction can occur between the amino group from the side chain of Lys residues and the C_{β} atom at the 5- or 6-position of BQ, resulting in covalent cross-linking.42 To investigate the relevance of 3-pyrazolone-fused BQ in modulating the reactivities with metal-free and metal-associated AB, we blocked the C_a and C₆ atoms by fusing a benzene group, creating 3-pyrazolonefused 1,4-naphthoquinone (NQ) compounds. The annulation of BQ in compound 1 with the benzene ring produced compound 3, which allowed us to examine the role of the 3-pyrazolonefused BQ in regulating Aβ aggregation mediated by compound 1. Given that the presence of multiple aromatic rings in compounds enhances their ability to interact with Aβ, ^{22,43–49} we further designed compounds 4 and 5 to maintain the same number of aromatic rings as in compound 1 by eliminating the p-tolyl group at the 2-position. In the case of compound 5, the number of donor atoms for hydrogen bonding and metal binding necessary for interacting with metal-free and metalbound Aβ was preserved by including a methyl group at the 2position.

As summarized in Fig. 1c, we synthesized heterocycle-fused BQ compounds, 1 and 2a. Compound 1a was diazotized to generate a diazonium salt intermediate, which was then reduced with SnCl₂ to synthesize a hydrazine intermediate. Intramolecular cyclization under acidic conditions produced an indazolone compound, 1b.50 N-Benzylation of 1b with benzyl chloride under basic conditions yielded an N-benzylated compound, 1c.51 Next, 1d was generated via a Cu(II)-catalyzed Chan-Evans-Lam coupling reaction of 1c with p-tolylboronic acid.52 The removal of the methyl moiety in the methoxy group of 1d using boron tribromide (BBr₃) afforded compound 1e.⁵³ Oxidation of 1e with ceric ammonium nitrate led to the synthesis of 2a.54,55 Finally, compound 1 was obtained by the Pdcatalyzed hydrogenolysis of 2a in a quantitative yield.56,57 The synthetic routes and methods for heterocycle-fused NQ compounds (3-5) have been recently described in the patent literature.58 The characterization of the compounds is summarized in Fig. S1-S5.†

Effects of compounds on the aggregation of metal-free $A\beta$ and metal- $A\beta$

To determine the impact of our compounds on the aggregation of metal-free and metal-bound A β , we initially monitored the size distribution of the resultant metal-free A β and metal-A β species upon incubation with the compounds by gel electrophoresis with western blotting (gel/western blot) using an anti-A β antibody (6E10). In the inhibition studies presented in

Fig. 2a and S6,† A β peptides were incubated with and without the compounds, both in the absence and presence of metal ions, to assess their effects on the formation of A β aggregates [Fig. 2a(i)]. As displayed in Fig. 2a(ii), heterocycle-fused BQ compounds, particularly 2a, exhibited significant reactivity towards A β aggregation in both the absence and presence of metal ions.

Under metal-free conditions, treatment of 1 with $A\beta_{42}$ decreased the intensity of smearing at 5-11 kDa. Notably, 2a intensified smearing across the gel lane from 5 kDa to 245 kDa. On the other hand, 3 and 4 moderately increased the band intensity at 48-135 kDa, and 5 diminished the signal intensity below ca. 20 kDa. In the presence of Cu(II), the intensity of bands at 35-135 kDa increased upon treatment with 1. The Cu(II)-Aβ₄₂ sample treated with 2a showed an increase in the signal intensity at 5-135 kDa. For Zn(II)-Aβ₄₂, the samples added with heterocycle-fused BQ compounds showed a clearer band at ca. 35 kDa relative to the compound-untreated sample, and 2a additionally intensified the smearing at 5-20 kDa. Regarding heterocycle-fused NQ compounds, the band intensity for metal-A β_{42} species at *ca*. 35 kDa was slightly enhanced by 3. The signal at the same size was also darkened for $Zn(II)-A\beta_{42}$ treated with 4. A significant change in the size distribution of metal-A β_{42} was not observed with the addition of 5. The results of the inhibition experiments with $A\beta_{40}$ also demonstrated the noticeable ability of heterocycle-fused BQ compounds, especially 2a, to alter their aggregation pathways. Compound 2a significantly changed the size distribution of metal-free $A\beta_{40}$, producing smearing between 5 kDa and 35 kDa, whereas 1 and 3-5 did not induce notable shifts in the size distribution of metal-free $A\beta_{40}$. Under Cu(II)-present conditions, new discrete bands at 5-11 kDa were detected when 1 and 2a were incubated with A β_{40} . Conversely, 3–5 did not modify the size distribution of the aggregates. For the sample of Zn(II)-A β_{40} incubated with 2a, explicit bands between 5 kDa and 11 kDa were revealed, while relatively minor reactivity was observed for the samples upon addition of 1 and 3-5, exhibiting blurred bands right above 5 kDa.

As depicted in Fig. 2a(iii), we further investigated the morphologies of the resultant AB aggregates by transmission electron microscopy (TEM), which can image large Aß aggregates that have difficulty in penetrating the gel matrix.25 Under metal-untreated conditions, all compounds fragmented long fibrils, shown in compound-free $A\beta_{42}$, into shorter filamentous species. When heterocycle-fused BQ compounds were incubated with $Cu(\pi)$ -A β_{42} , they produced dramatically smaller aggregates with amorphous qualities compared to the compound-untreated sample, which contained larger aggregates. Treatment with 3 also reduced the size of $A\beta_{42}$ aggregates with Cu(II), but 4 and 5 did not result in considerable morphological and size changes in Cu(π)-Aβ₄₂ aggregates. In the presence of Zn(II), all compounds altered the morphologies of Zn(II)-Aβ₄₂ aggregates from fibrillary to less structured species. In particular, the 2a-treated sample manifested chopped fibrils with small-sized amorphous aggregates. For the samples of $A\beta_{40}$, a mixture of chopped fibrils with unstructured aggregates was detected by addition of 2a under metal-free

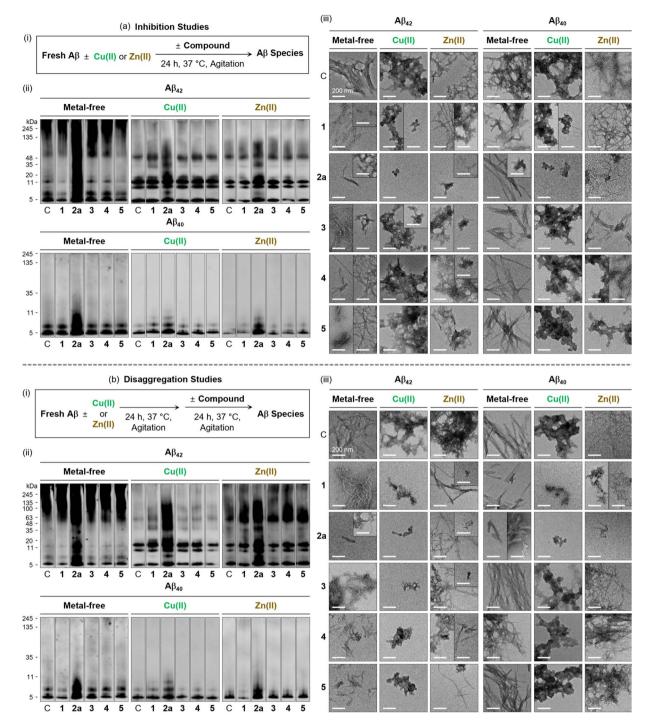


Fig. 2 Effects of compounds on the aggregation of metal-free and metal-treated Aβ analyzed by two types of experiments [(a) inhibition and (b) disaggregation studies]. (i) Scheme of Aß aggregation experiments. (ii) Size distribution of the resultant Aß species probed by gel/western blot using an anti-A β antibody (6E10). Lane C: A β \pm Cu(II) or Zn(II). The original gel images are shown in Fig. S6.† (iii) Morphologies of the resultant A β species monitored by TEM. Conditions: $[A\beta] = 25 \mu M$; $[Cu(III) \text{ or } Zn(III)] = 25 \mu M$; $[compound] = 25 \mu M$ (1% v/v DMSO); 20 mM HEPES, pH 7.4 [for metal-free and Zn(ii)-added samples] or pH 6.8 [for Cu(ii)-added samples], 150 mM NaCl; 37 °C; 24 h; constant agitation. Scale bar = 200 nm.

conditions, whereas compound-free, 1-, and 3-5-treated samples showed primarily fibrillary species. In the presence of Cu(II), heterocycle-fused BQ compounds generated smaller amorphous aggregates compared to compound-untreated and heterocycle-fused **NQ** compound-added Cu(II)-A β_{40} . In the case of Zn(II)-A β_{40} samples, the formation of thinner filamentous

species, without and with smaller less-structured aggregates, was induced by 1 and 2a, respectively, different from the entangled fibrillary species detected in the compound-free sample. Heterocycle-fused NQ compounds facilitated aggregation into less organized structures and generated chopped fibrils in 3-incubated Zn(II)-A β_{40} . Collectively, both heterocyclefused BQ and NQ compounds alter the formation of $A\beta$ aggregates to different degrees.

Moreover, the disaggregation studies were performed to identify the ability of the compounds to disassemble preformed Aß aggregates or alter their aggregation pathways further, as illustrated in Fig. 2b and S6.† The samples were prepared by pre-incubating AB with and without metal ions to form preformed AB assembles, followed by treatment with the compounds [Fig. 2b(i)]. As shown in Fig. 2b(ii), both heterocycle-fused BQ and NQ compounds were capable of shifting the size distribution of $A\beta_{42}$ aggregates under metalfree and metal-present conditions. All compounds increased the intensity of bands at 48-63 kDa without metal ions, with 2a prominently varying the size distribution of metal-free Aβ₄₂ species at 5-63 kDa. In the case of metal- $A\beta_{42}$ incubated with heterocycle-fused BQ compounds, the signal intensity was enhanced at 20-100 kDa and 5-135 kDa following treatment with 1 and 2a, respectively. Heterocycle-fused NQ compounds slightly elevated the level of Cu(II)-Aβ₄₂ species at 35–100 kDa, and 5 also decreased the signal at 5-11 kDa. In the presence of Zn(II), the size distribution of A β_{42} at 63–135 kDa was shifted by 3-5. For the experiments with $A\beta_{40}$, incubation of heterocyclefused **BQ** compounds remarkably altered the amount of $A\beta_{40}$ oligomers between 5 kDa and 11 kDa under both metal-free and metal-present conditions, compared to the samples without compounds and those treated with heterocycle-fused NQ compounds, which showed less reactivity.

In the TEM studies depicted in Fig. 2b(iii), 1 and 3 produced thinner metal-free $A\beta_{42}$ fibrils, compared to relatively thicker and longer fibrillary aggregates observed in the compound-free sample, while 2a, 4, and 5 shortened the fibrils. The size of $Cu(\pi)$ -A β_{42} aggregates was dramatically reduced by addition of both heterocycle-fused **BQ** and **NQ** compounds. Under Zn(II)treated conditions, all compounds generated thinner filamentous species different from the bundle of fibrillar aggregates of compound-free Zn(II)-Aβ₄₂, and smaller amorphous aggregates were also shown for the samples incubated with 1, 2a, and 3. For $A\beta_{40}$, both 1 and 2a remarkably truncated fibrillary species displayed in the compound-free sample under metal-free conditions. In the presence of Cu(II), the addition of heterocycle-fused BQ compounds induced the formation of smaller-sized amorphous aggregates. Under Zn(II)-present conditions, relative to Zn(II)-A β_{40} aggregates with fibrillary characteristics, a mixture of amorphous and fibrillary aggregates was monitored upon treatment of 1 with $A\beta_{40}$ aggregates, and the addition of 2a triggered the formation of chopped fibrils. In contrast, little or no morphological alteration was detected for the samples prepared with heterocycle-fused NQ compounds. Overall, both heterocycle-fused BQ and NQ compounds can disassemble the preformed AB aggregates or additionally alter their aggregation pathways in the absence and presence of metal ions to distinct extents. Heterocycle-fused BQ compounds exhibited significant modulatory effects on the aggregation of metal-free and metal-associated AB over heterocycle-fused NO compounds. This indicates that the structural modifications on the BQ moiety, without the fusion of an additional benzene ring at the 5- and 6-positions, are

essential for impacting the assembly of $A\beta$ with and without metal ions. Notably, the enhanced ability of 2a relative to 1 in changing the size distribution and morphology of $A\beta$ aggregates under both metal-free and metal-treated conditions suggests that the benzyl ring at the 1-position of 2a contributes to its regulatory reactivity with metal-free $A\beta$ and metal- $A\beta$.

Mechanistic studies (I): interactions of compounds with metal-free $A\beta$

To elucidate the distinct reactivities of heterocycle-fused BQ and NQ compounds towards AB aggregation, we investigated their interactions with A β peptides. Considering that A β_{42} is more pathogenic than Aβ₄₀,⁵⁹ and the modulative effects of our compounds were more discernible in the samples with AB₄₂ rather than $A\beta_{40}$ (Fig. 2), we focused on analyzing the direct interactions between the compounds and $A\beta_{42}$ by electrospray ionization-mass spectrometry (ESI-MS). As presented in Fig. 3a, metal-free $A\beta_{42}$ incubated with 1 and 3-5 showed the peaks assigned to non-covalent adducts of $A\beta_{42}$ with the compounds at 1589 m/z, 1606 m/z, 1576 m/z, and 1580 m/z, respectively. In the presence of 2a (344 Da), a peak corresponding to the addition of 342 Da to the 3+-charged A β_{42} monomer (4511 Da) was detected at 1619 m/z, likely representing the covalent adduct of 2a with metal-free A β_{42} . Further fragmental analysis of the peak through collision-induced dissociation (CID) identified the amino acid residue in $A\beta_{42}$ involved in the formation of this covalent cross-linking. In tandem MS (ESI-MS²) depicted in Fig. 3b, b fragments smaller than b_{16} were observed in their native forms. Starting from b_{16} , the b fragments were detected with and without the addition of 342 Da, suggesting Lys16 as the amino acid residue forming a covalent bond with 2a. Native forms of b fragments were no longer monitored beyond b_{27} , indicating that Lys28 could also be a site for covalent crosslinking with 2a.

Lys residues in several amyloidogenic peptides and proteins have been reported to facilitate both hydrophobic and electrostatic interactions via flexible butyl and ϵ -ammonium groups in their side chains, respectively, in the intramolecular and intermolecular manner. ^{60,61} More specifically, Lys16, adjacent to the self-recognition site (Fig. 1a), forms a salt bridge with Glu22 in the fibril structure of Aβ⁶² or is exposed to the solvent, making it available for interacting with other monomers. ⁶³ Additionally, Lys28 stabilizes the turn structure of monomeric Aβ through a salt bridge with Asp23, which can mediate the folding and assembly of Aβ. ^{62,64–66} Therefore, the covalent bond formation between the **BQ** moiety in **2a** and Lys residues of Aβ₄₂ could be critical in altering its aggregation pathways.

Interestingly, this phenomenon was observed only in the sample incubated with 2a, not with 1. This implies the involvement of the benzyl ring at the 1-position of 2a in the covalent adduct formation with $A\beta_{42}$. Based on the potential interactions between 2a and monomeric $A\beta_{42}$ visualized by docking studies (Fig. 3c), the benzyl group can participate in CH- π interactions with Lys16 and Phe20. This suggests that this benzyl functionality may facilitate the recruitment of 2a by $A\beta_{42}$, followed by the cross-linking reaction between them.

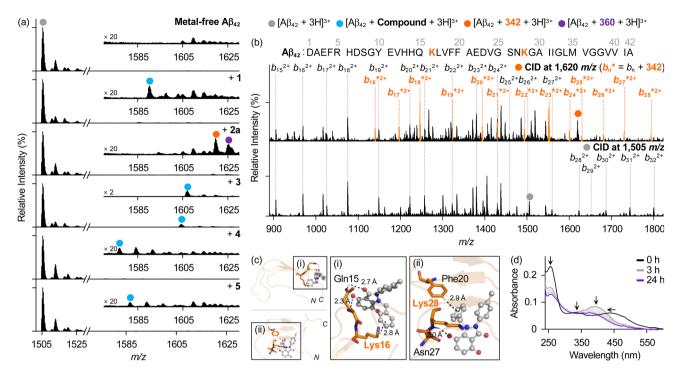


Fig. 3 Interaction of compounds with metal-free A β_{42} and 2a's chemical transformation. (a) ESI-MS studies of metal-free A β_{42} with and without compounds. Conditions: $[A\beta_{42}] = 25 \mu M$; $[compound] = 25 \mu M$ (1% v/v DMSO); 20 mM ammonium acetate, pH 7.4; 37 °C; 24 h; constant agitation. The samples were diluted 25-fold with H_2O before injection into the mass spectrometer. (b) ESI-MS² investigations of metal-free $A\beta_{42}$ with (top; 1620 m/z) and without (bottom; 1505 m/z) covalent bond formation with 2a. (c) Possible interactions of 2a with metal-free A β_{42} (xx98156 (ref. 26)) visualized by docking studies. Nine docked models were obtained with binding energies ranging from $-5.9 \text{ to } -5.5 \text{ kcal mol}^{-1}$ and two representative models are indicated in (i) and (ii). (d) Chemical transformation of 2a monitored by Abs spectroscopy. Conditions: [2a] = 25 μ M (1% v/v DMSO); 20 mM HEPES, pH 7.4, 150 mM NaCl; 37 °C; 24 h; no agitation.

Furthermore, 2a is positioned near the self-recognition and Cterminal regions, encompassing the Lys16 and Lys28 residues as Michael donors for forming covalent bonds with 2a. Multiple hydrogen bonds were also predicted for the carbonyl O atoms located at the 3-, 4-, and 7-positions of 2a with the amide group from the side chain of Gln15 and the backbone amide moieties between Gln15 and Lys16, and Asn27 and Lys28.

Moreover, an additional peak at $1624 \, m/z$, indicating a mass shift of +360 Da from metal-free $A\beta_{42}$ monomer, appeared upon incubation with 2a. The nature of this peak was first analyzed by examining the chemical transformation of 2a by electronic absorption (Abs) spectroscopy, as shown in Fig. 3d. The optical spectrum of 2a in aqueous solution exhibited three Abs bands at ca. 255 nm, 330 nm, and 440 nm, consistent with those observed for BQ.67-69 After 24 h of incubation, the spectral features at the shortest wavelength diminished, while a hypsochromic shift from ca. 440 nm to ca. 390 nm and a bathochromic shift from ca. 255 nm to ca. 265 nm were monitored, indicative of o-hydroxylation of the BQ moiety in 2a.67-69 Previous studies have proposed that the reaction of BQ with H₂O at neutral pH triggers the formation of enolized BQ [4,6dihydroxycyclohexa-2,4-dien-1-one; BQ(H2O)], which can be rearranged into 1,2,4-trihydroxybenzene (THB) (Fig. S7a†).67-69 Oxidation of BQ(H₂O) by BQ and auto-oxidation of THB subsequently produce 2-hydroxy-BQ (2-hydroxycyclohexa-2,5diene-1,4-dione).67-69 As depicted in Fig. S8,† the mass spectrum of 2a incubated for 24 h in aqueous solution displayed a peak

assigned to [2a + 16 + H]⁺, plausibly denoting 2a substituted with a hydroxyl group (2b; vide infra). Thus, the peak at 1624 m/z in the mass spectrum of 2a-treated metal-free $A\beta_{42}$ (Fig. 3a) can arise from the non-covalent adduct formation between monomeric $A\beta_{42}$ and **2a** containing a hydroxylated **BQ** moiety.

In summary, both heterocycle-fused BQ and NQ compounds can directly interact with monomeric $A\beta_{42}$. We further confirmed that 2a, compared to 1, can participate in the generation of covalent bonds with $A\beta_{42}$. These suggest that a supplementary benzyl moiety at the 1-position of 2a is required for sufficient contact with Aβ, consequently producing a covalent cross-link to the peptide via BQ. Moreover, 2a undergoes o-hydroxylation of the BQ moiety and, subsequently, directly interacts with A\beta. Such 2a-driven interactions with metal-free AB support its significant reactivity towards AB in both the absence and presence of metal ions.

Mechanistic studies (II): interactions of chemically transformed 2a with metal ions, metal-free AB, and metal-AB

To gain a better understanding of the involvement of chemically transformed 2a in its reactivity towards the aggregation of metal-free A β and metal-A β , we synthesized compound 2b, which is 2a with a hydroxyl group at the 6-position, and analyzed its interactions with metal ions, metal-free AB, and metal-Aβ. As presented in Fig. 4a, compound 1e was dibrominated with Br2, yielding 4,6-dibromo-indazolone and the

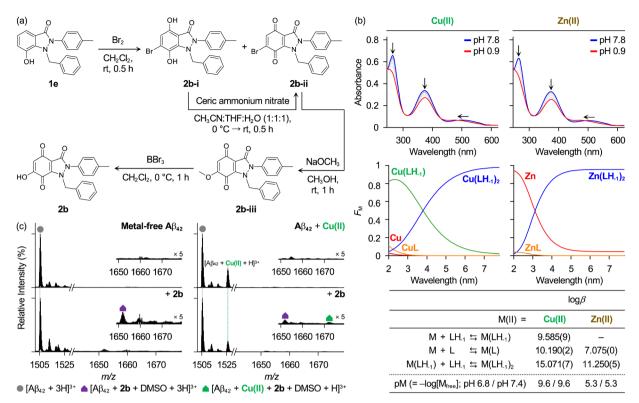


Fig. 4 Preparation of **2b** and its ability to interact with metal ions, metal-free $A\beta_{42}$, and $Cu(II)-A\beta_{42}$. (a) Synthetic routes to **2b** [1-benzyl-6-hydroxy-2-(p-tolyl)-1H-indazole-3,4,7(2H)-trione]. (b) Solution speciation studies of metal-**2b** (L) complexes. Abs variable-pH titration spectra (top) and the solution speciation diagram (middle) were obtained by titrations of metal-ligand complexes (F_M = fraction of species at a given pH). The log β and pM values of metal-ligand complexes are summarized in the table (bottom). The errors in the last digit are shown in parentheses. Charges are omitted for clarity. Conditions: [**2b**] = 50 μ M; [Cu(II)] or Zn(III)] = 25 μ M; room temperature; I = 0.1 M NaCl. (c) Interactions of **2b** with metal-free $A\beta_{42}$ and $Cu(III)-A\beta_{42}$ monitored by ESI-MS. Conditions: [$A\beta_{42}$] = 25 μ M; [Cu(III)] = 25 μ M; [**2b**] = 25 μ M (1% v/v DMSO); 20 mM ammonium acetate, pH 7.4 (for metal-free samples) or pH 6.8 [for Cu(III)-treated samples]; 37 °C; 24 h; constant agitation. The samples were diluted 25-fold with H₂O before injection into the mass spectrometer.

unexpected 6-bromo-benzoquinone (6-Br-BQ) derivative. 70 We assumed that the latter was oxidized by H2O and Br2 derived from 4,6-dibromo-indazolone. Consequently, the reaction conditions were established to synthesize 6-bromo-4,7-dihydroxyindazolone, 2b-i, and 6-Br-BQ derivative, 2b-ii. 2b-i was oxidized to 2b-ii using ceric ammonium nitrate.54,55 The subsequent methoxylation of **2b-ii** with sodium methoxide (NaOCH₃) resulted in the formation of a 6-methoxy-benzoquinone derivative, 2b-iii. 71 Finally, 2b was obtained by the demethylation of 2b-iii using BBr₃ in a high yield. 72 The characterization of 2b is summarized in Fig. S9.† It is noteworthy that spectral features of 2b in aqueous solution are similar to those of 2a incubated for 24 h, but with greater optical intensity. This observation supports that 2a undergoes chemical transformations into 2b and various other products, including less soluble aggregates (vide infra).67-69,73-75

o-Hydroxylation of BQ in 2a introduces an additional metalchelation site, composed of a pair of O donor atoms from hydroxyl and carbonyl groups at the 6- and 7-positions, respectively. To explore the capacity of 2b to bind metal ions in aqueous solution, solution speciation studies with Abs variablepH titrations were performed. Initially, the solution of 2b (L) was titrated with small aliquots of HCl under metal-free conditions to estimate its acidity constants (K_a). As illustrated in Fig. S10,† two p K_a values, 0.627(5) and 4.827(5), were calculated. The solution speciation diagram displayed the presence of anionic (LH₋₁), neutral (L), and monoprotonated (LH) forms of **2b** in the pH range from 2 to 9. The anionic form of the ligand was predicted to predominate at pH 6.8, a condition representing a physiologically acidotic environment where A β assembly can be accelerated in the presence of Cu(π), ⁷⁶ and at physiological pH 7.4, accounting for *ca.* 98% and *ca.* 100% relative abundance, respectively.

Subsequently, solution speciation experiments of 2b were conducted in the presence of $Cu(\pi)$ and $Zn(\pi)$ to identify the metal-to-ligand stoichiometry and ligand's metal-binding affinities. As shown in Fig. 4b, the stability constants (β) were calculated for the formation of metal-2b complexes plausibly present at pH 0.9–7.8. Metal complexes with the 1:2 metal-to-ligand stoichiometry were indicated as a major species at pH 6.8 and pH 7.4. Based on the protonation and metal complexation of 2b at the given pH and concentrations of the ligand and metal ions, the values of pM $(-\log[M_{free}],$ where $[M_{free}]$ denotes the concentration of unchelated metal ions) were computed to predict the relative metal-binding ability of 2b under our experimental conditions. 77,78 At pH 6.8 and pH 7.4, 2b presented

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pCu and pZn values of 7.7 and 5.3, respectively. It should be noted that the metal-binding affinity of 2a could not be experimentally determined due to its chemical transformation in aqueous solution. Instead, we evaluated the metal-binding ability of 6 which is derivatized from 3, which features the substitution of the amino group at the 1-position with a methyl group and the replacement of the p-tolyl group at the 2-position with a phenyl ring.58,79 This compound is relatively stable under our experimental conditions and has potential for metal chelation through two carbonyl O donor atoms in the 3- and 4positions, similar to 2a; however, no significant spectral change was detected in Abs spectra of the compound upon addition of Cu(II) (Fig. S11†). This corroborates the limited capability of 2a to interact with metal ions under our experimental settings. These results underline that the chemical transformation of 2a imparts the ability to chelate metal ions (Fig. 1b).

To verify the interaction between 2b and $A\beta_{42}$ with and without metal ions, ESI-MS studies were carried out. As depicted in Fig. 4c and S12,† new peaks at 1651 m/z were detected upon addition of **2b** to $A\beta_{42}$ under both metal-free and metal-present conditions, assigned to $[A\beta_{42} + 2\mathbf{b} + DMSO + 3H]^{3+}$. In the presence of Cu(II), a peak at 1671 m/z corresponding to $[A\beta_{42} + Cu(II) +$ 2b + DMSO + H]3+ was observed, indicating the formation of a ternary complex composed of $A\beta_{42}$, Cu(II), and **2b**. This compound also generated a ternary complex with Zn(II)-Aβ₄₂ (Fig. S12†). The pM values of 2b (Fig. 4b), comparable to those of A β (ca. 6-12 for pCu at pH 6.5-7.4; ca. 6 for pZn at pH 7.4), ^{11,13,80,81} support such competitive metal coordination of 2b with Aβ. These findings demonstrate that 2b can interact with metal ions as well as $A\beta$ in both the absence and presence of metal ions. Consequently, the chemical transformation of 2a, including ohydroxylation likely into 2b, enables it to chelate metal ions in a bidentate manner with binding affinities similar to those of Aβ, ultimately impacting the self-assembly of metal-Aβ.

Mechanistic studies (III): impact of 2b on the aggregation of metal-free $A\beta$ and metal-A β

The effect of **2b** on the assembly of metal-free Aβ and metal-Aβ was further probed by gel/western blot using 6E10 and TEM. In the inhibition studies, the size distribution of Aβ species with and without metal ions was not significantly influenced by 2b, as delineated in Fig. 5a and S6;† however, the impact of 2b on the aggregation of the peptides in the absence and presence of metal ions was discernibly shown in the TEM studies. The resultant $A\beta_{42}$ and $A\beta_{40}$ aggregates produced in the presence of 2b were analyzed as shorter fibrils relative to compound-free samples, which showed significant fibrillar aggregates under metal-free and Zn(II)-added conditions, as illustrated in Fig. 5b. Upon incubation of 2b with Cu(II)-Aβ, a mixture of smaller aggregates with amorphous and fibrillary characteristics was generated for $A\beta_{42}$ and $A\beta_{40}$, differing from compound-free samples, which mainly displayed larger aggregates. Together, 2b and 2a affect the formation of Aβ aggregates in the absence and presence of metal ions to distinct extents.

In the disaggregation studies conducted by gel/western blot using 6E10, as shown in Fig. S6 and S13, \dagger the reactivity of 2b

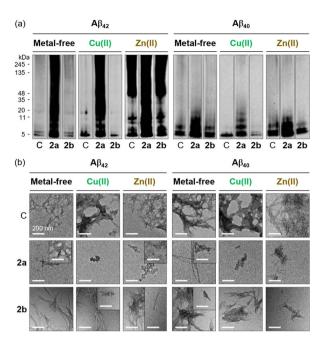


Fig. 5 Impact of **2a** and **2b** on the formation of metal-free and metal-associated A β aggregates. (a) Size distribution of the resultant A β species analyzed by gel/western blot using 6E10. Lane C: A β \pm Cu(II) or Zn(III). The original gel images are shown in Fig. S6.† (b) Morphologies of the resultant A β species investigated by TEM. Conditions: [A β] = 25 μ M; [Cu(II) or Zn(III)] = 25 μ M; [compound] = 25 μ M (1% v/v DMSO); 20 mM HEPES, pH 7.4 [for metal-free and Zn(III)-containing samples] or pH 6.8 [for Cu(III)-containing samples], 150 mM NaCl; 37 °C; 24 h; constant agitation. Scale bar = 200 nm.

with metal-free and metal-associated AB was not clearly monitored. TEM, however, presented that 2b-added $A\beta_{42}$ and $A\beta_{40}$ samples included shorter and chopped fibrils, different from compound-unadded samples which exhibited bigger aggregates, under metal-free and metal-added conditions. These findings indicate that both 2b and 2a can disassemble preformed Aβ aggregates or additionally alter their aggregation profiles in the absence and presence of metal ions to varying degrees. Overall, 2b, as the chemically transformed form of 2a, can modulate the aggregation pathways of AB in both the absence and presence of metal ions, potentially contributing to the reactivity of 2a with metal-free Aβ and metal-Aβ. Given that Aβ samples incubated with 2a contain various chemically transformed forms of 2a, the reactivity of 2a towards AB aggregation with and without metal ions may be influenced not only by 2b but also by other chemically transformed forms, as well as intact 2a (Fig. S7b†). It should be noted that we monitored the formation of 2b' upon incubation of 2a for 6 h in aqueous media, followed by its disappearance after 24 h, as depicted in Fig. S8.† According to previous studies, compounds with multiple aromatic rings can participate in π - π stacking, leading to the generation of relatively less soluble aggregates that become optically undetectable.71,72 The aggregates present in 2aand 2b-incubated samples, which differ in composition and relative abundance, may further influence the extent to which 2a and 2b affect the assembly of Aβ. In addition, as shown in

Fig. S7c, \dagger a hydroxyl group at the 6-position of **2b** enables additional hydrogen bonding with A β , potentially altering its reactivity with A β in the absence and presence of metal ions compared to **2a**.

Mechanistic studies (IV): metal-binding properties of compounds and modifications of A β by compounds in the presence of Cu(II)

Apart from **2a**, **1** and **3–5** share two possible metal-chelation sites: N and O donor atoms at the 1- and 7-positions, respectively, and two O donor atoms at the 3- and 4-positions. Given that their metal-binding ability can mediate their reactivity with metal-A β , metal-binding properties of **1** as a representative molecule were analyzed. Job's method of continuous variation was used to determine the Cu(II)-to-ligand stoichiometry (Fig. S14a†). We mixed **1** and Cu(II) at various ratios while maintaining their total concentration and traced the Abs at 250 nm. The Job plot revealed a maximum intensity at a mole fraction of Cu(II) of *ca*. 0.3 with an asymmetrical curve shape, suggesting that a mixture of Cu(II)-**1** complexes with 1: 1 and 1: 2 Cu(II)-to-ligand ratios could exist in aqueous solution. ^{82,83} Next, we titrated the aqueous solution of **1** with Cu(II) and

monitored the spectral changes at 250 nm (Fig. S14b and c†). These led to estimating the dissociation constants (K_d) of **1** for Cu(II) to be in the micromolar range. Although this Cu(II)-binding affinity is likely insufficient to effectively interact with the primary Cu(II)-binding site in A β , which has a nanomolar affinity, the presence of a second Cu(II)-binding site, with an affinity in the micromolar range, ^{10,30,84} implies that the interaction of **1** and **3**–5 with Cu(II) at this secondary site in A β may still be feasible. It should be noted that the p K_a value of **1** and its Zn(II)-binding affinity could not be determined due to the limited spectral change within the pH range monitored for our measurements without dilution effects and upon incubation with Zn(II), respectively, under our experimental conditions (Fig. S15†).

We then examined the samples of metal-A β_{42} incubated with heterocycle-fused **BQ** or **NQ** compounds by ESI-MS. As presented in Fig. 6a, **1** and 3–5 formed non-covalent adducts with A β_{42} in the presence of Cu(II), similar to what was observed under metal-free conditions (1589 m/z, 1606 m/z, 1576 m/z, and 1580 m/z for **1**, **3**, **4**, and **5**, respectively). **3** additionally indicated the ternary complexation with A β_{42} and Cu(II) (1626 m/z). Under Zn(II)-present conditions, all compounds could also interact with A β_{42} similar to their behavior in the absence of metal ions

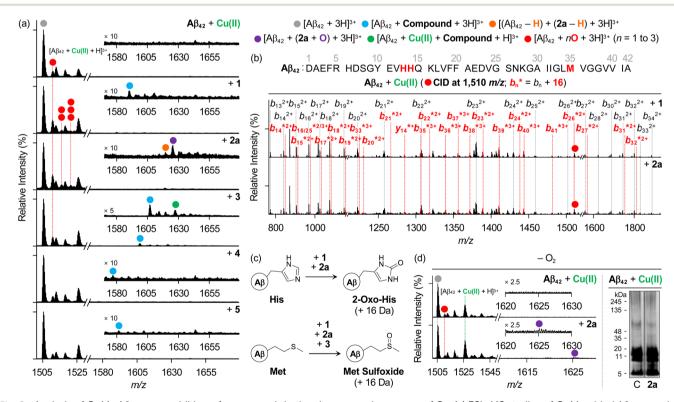


Fig. 6 Analysis of $Cu(II)-A\beta_{42}$ upon addition of compounds in the absence and presence of O_2 . (a) ESI-MS studies of Cu(II)-added $A\beta_{42}$ samples with and without compounds under aerobic conditions. The number of red circles represents that of O atoms incorporated into the $A\beta_{42}$ monomer. Conditions: $[A\beta_{42}] = 25 \mu M$; $[Cu(II)] = 25 \mu M$; $[Cu(III)] = 25 \mu M$; $[Cu(IIII)] = 25 \mu M$; $[Cu(IIIIII)] = 25 \mu M$; $[Cu(IIIII)] = 25 \mu M$; $[Cu(IIIII]] = 25 \mu M$; $[Cu(IIII]] = 25 \mu M$; [Cu(IIII

(Fig. S16†). Interestingly, consecutive peaks appeared in the $Cu(\pi)$ -A β_{42} samples added with heterocycle-fused **BQ** compounds at 1510 m/z, 1515 m/z, and 1520 m/z, corresponding to singly, doubly, and triply oxidized $A\beta_{42}$, respectively. Upon applying the CID energy to the peak at 1510 m/z, as depicted in Fig. 6b, b fragments from b_{13} were exhibited in their nonoxidized and oxidized forms, indicating that heterocycle-fused BQ compounds oxidized His13, His14, or Met35. The CID of the peaks assigned to doubly and triply oxidized peptides could not be well-resolved under our experimental settings, but three plausible oxidation sites imply that heterocycle-fused BQ compounds simultaneously oxidize two or three of His13, His14, and Met35.

As summarized in Fig. 6c, the structural conversion of His into 2-oxo-His, caused by heterocycle-fused BQ compounds, was reported to change the polarity of His, leading to disturbances in electrostatic interactions of AB peptides, thereby modifying their assembly profiles. 12,85 Additionally, His residues are included in the metal-coordination site in AB (Fig. 1a) and, thus, their structural modification affects metal binding to AB peptides and their aggregation pathways in the presence of metal ions.11 Furthermore, the oxidative modification of the Met residue results in Met sulfoxide, which can alter the secondary structures across the peptides, reducing their hydrophobic and electrostatic associations.86-89 In the case of the 3-treated sample, a singly oxidized Aβ₄₂ monomer was also shown (Fig. 6a), and Met35 was revealed to be the oxidation site, as detected for the compound-free Cu(II)-Aβ₄₂ sample (Fig. S17†).

To verify the involvement of copper-O2 chemistry in the oxidative transformation of Aβ, 90,91 the samples of Cu(II)-Aβ₄₂ incubated with and without 2a under anaerobic conditions were further explored, as shown in Fig. 6d. Distinct from the mass spectra obtained under aerobic conditions (Fig. 6a), there was no notable variation in the signal intensity of the singly oxidized $A\beta_{42}$ monomer upon treatment with 2a, and the peaks assigned to multiply oxidized peptides were not observed in the absence of O2. Thus, copper-O2 chemistry may promote the oxidative modifications of A β in the presence of 2a. Notably, A β_{42} still formed a non-covalent adduct with the chemically transformed 2a (2a containing o-hydroxylated BQ) under anaerobic conditions. In the gel/western blot using 6E10, 2a produced Aβ oligomers of 5-11 kDa. According to the ESI-MS results, Cu(II)added $A\beta_{42}$ aggregation could be varied even without O_2 by the chemical transformation of 2a and its subsequent interaction with $A\beta_{42}$. Our biophysical and biochemical observations substantiate the direct interactions of both heterocycle-fused **BO** and **NO** compounds with Aβ in the presence of metal ions. Heterocycle-fused BQ compounds, particularly 2a, can concurrently oxidize multiple amino acid residues in AB in the presence of $Cu(\pi)$ depending on the availability of O_2 , changing the aggregation patterns of Cu(II)-A β with and without O_2 in a distinct manner.

Cytotoxicity of compounds and their influence on A\(\beta\)-induced cytotoxicity with and without metal ions

As our compounds affect the aggregation pathways of Aβ in the absence and presence of metal ions (Fig. 2, 5 and S13†), their ability to control the cytotoxicity induced by AB with and without metal ions was assessed by the MTT assay [MTT = 3](4,5-dimethylthiazol-2-yl)-2,5-diphenyltetrazolium employing human neuroblastoma SH-SY5Y (5Y) cells (Fig. 7a). Prior to conducting cell studies with AB, the cytotoxicity of heterocycle-fused BQ and NQ compounds and 2b was evaluated under metal-free and metal-present conditions and compared

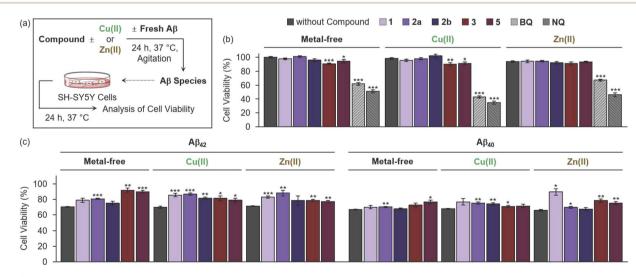


Fig. 7 Cytotoxicity of compounds under metal-free and metal-treated conditions and their impact on the cytotoxicity induced by metal-free and metal-associated Aβ. (a) Scheme of the experiments. (b) Viability (%) of 5Y cells upon incubation with heterocycle-fused BQ or NQ compounds, 2b, BQ, or NQ with and without metal ions. (c) Influence of heterocycle-fused BQ and NQ compounds and 2b on the cytotoxicity mediated by metal-free and metal-treated Aβ. The MTT assay provided the cell survival (%), calculated relative to that of cells treated with an equivalent amount of the buffered solution (20 mM HEPES, pH 7.4 or pH 6.8, 150 mM NaCl) with 0.2% v/v DMSO. Conditions: $[A\beta] = 25 \mu M$; [Cu(II)]or $Zn(III) = 25 \mu M$; [compound] = 25 μM (0.2% v/v DMSO); 20 mM HEPES, pH 7.4 [for metal-free and Zn(III)-added samples] or pH 6.8 [for Zn(III)-added samples added samples], 150 mM NaCl. All values are indicated as mean \pm SEM. Statistically significant difference from compound-untreated controls was determined by a two-sided unpaired Student's t-test (*P < 0.05, **P < 0.01, or ***P < 0.001).

with that of **BQ** and **NQ**. As shown in Fig. 7b, no significant cytotoxicity was observed for heterocycle-fused **BQ** and **NQ** compounds and **2b** in both the absence and presence of metal ions (cell survival over 90%) at the concentrations used for cell studies with Aβ. In contrast, **BQ** and **NQ** severely decreased cell survival by *ca*. 33–65% regardless of the presence of metal ions. These results demonstrate that the fusion of **BQ** and **NQ** on the 3-pyrazole framework, with and without the substitution at the 1- or 2-position or both, as well as the 6-position, can mitigate their cytotoxicity. It should be noted that the cytotoxicity of 4 could not be determined due to its limited solubility under our experimental conditions.

Moving forward, Aβ species preincubated with and without heterocycle-fused BQ or NQ compounds in the absence and presence of metal ions were introduced into the cells. As illustrated in Fig. 7c, some compounds exhibited substantial mitigating effects on the cytotoxicity triggered by AB with and without metal ions. The amelioration of metal-free $A\beta_{42}$ induced cytotoxicity was achieved by addition of 2a, enhancing cell viability by ca. 10%. When 1 and 2a were incubated with metal-A β_{42} , the cell survival increased by ca. 12-17%. Treatment of **2b** decreased the cytotoxicity of Cu(II)-A β_{42} by *ca.* 12%. Cells added with metal-free and metal-associated $A\beta_{42}$ with heterocycle-fused NQ compounds also showed an increase in cell viability by ca. 6-22%, compared to those treated with compound-free samples. In the case of $A\beta_{40}$, the cell survival was enhanced upon treatment with 1 in the presence of Zn(II) and 2a with and without metal ions by ca. 23% and ca. 3-7%, respectively. The addition of **2b** with Cu(II)-A β_{40} improved cell viability by ca. 6%. Regarding heterocycle-fused NQ compounds, 3 lowered metal-A β_{40} -induced cytotoxicity by *ca.* 3-12%, and 5 also reduced cellular death mediated by $A\beta_{40}$ under metal-free and $Zn(\pi)$ -present conditions by ca. 9–10%. Overall, our investigations confirm that heterocycle-fused BQ and NQ compounds and 2b exert cytoprotective effects against metalfree and metal-associated Aβ.

Conclusions

AD is linked to the self-assembly of A β in the absence and presence of metal ions, resulting in neurotoxic aggregates observed as deposits in the patients' brains. ^{10–14,19,21,92} In an effort to modulate the aggregation of A β , small molecules that can be converted into BQ have been reported. ^{25,29,30,73} Following their chemical transformation, BQ forms covalent adducts with A β peptides, with a consequent impact on their assembly routes; ^{25,29,30} however, the inability to chelate metal ions and the significant toxicity of BQ have limited its reactivity with metal–A β and its biological applications. In this work, we demonstrate how to effectively design heterocycle-fused BQ compounds that modulate the aggregation of both metal-free A β and metal–A β by leveraging the strength of BQ (*e.g.*, covalent cross-linking to A β) while addressing its weaknesses (*e.g.*, poor metal chelation and toxicity).

We rationally designed bidentate ligands by combining **BQ** with the 3-pyrazolone framework, varying the type, position, and number of substituents on the 3-pyrazolone group without

(1 and 2a) and with (3-5) the annulation of BQ with a benzene ring. Our multidisciplinary investigations reveal that heterocycle-fused BQ compounds significantly regulate the aggregation of metal-free Aβ and metal-Aβ, while heterocyclefused NQ compounds lack this reactivity. These findings underscore the importance of the structural tuning and refinement of the BQ moiety in directing the ability of heterocycle-fused BQ compounds to change the self-assembly pathways of both metal-free Aβ and metal-Aβ. More importantly, 2a can covalently cross-link to AB via BQ with the assistance of the benzyl group at the 1-position for A β interaction. The BQ functionality of 2a further undergoes o-hydroxylation, enhancing its ability to chelate metal ions and, consequently, form a ternary complex with metal-Aβ. Moreover, heterocyclefused BO compounds simultaneously oxidize multiple amino acid residues in Aβ peptides added with Cu(II) and modify their assembly profiles in both the absence and presence of O₂ to varying degrees. Our compounds ultimately reduce the toxicity induced by AB with and without metal ions in living cells. Collectively, our findings highlight the structural variations and optimizations of BQ for designing chemical tools to target and control the aggregation and toxicity of both metal-free AB and metal-Aβ, as well as developing therapeutic agents to combat metal-entangled pathologies in amyloid diseases. In the near future, detailed structure-activity relationship studies of our compounds will be conducted to identify compounds with the highest efficiency and enhanced properties for biological applications, including metabolic stability and blood-brain barrier permeability, and proceed with their in vivo evaluations.

Data availability

All experimental details and data supporting the findings of this study are available within the paper and its ESI.† The data are also available from the corresponding authors upon reasonable request.

Author contributions

Y. Y., H. K., and M. H. L. designed the research. Y. Y. performed gel/western blot, TEM, ESI–MS, docking simulations, and solution speciation studies using Abs spectroscopy and analyzed the data. Y. Y. also conducted cell studies using the MTT assay. K. K. and H. K. synthesized all compounds tested in this work. Y. Y., K. K., H. K., and M. H. L. wrote the manuscript.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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