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Mechanistic insights into N₂ activation by RhCo₃ via d–d orbital coupling

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The synthesis of ammonia (NH₃) from nitrogen (N₂) under mild conditions is a great challenge, in which the electron-donating ability of the catalyst is the key for N₂ activation. In this work, the above process was studied using quantum chemical calculations using the density functional method. The results show that Rh and Co exhibit unsaturated d-electron configurations, with d–d orbital coupling occurring within the RhCo₃ metal cluster, resulting in bimetallic synergy and spin effects. The Rh atom serves as an electron modulation center and active site for reactant activation, while the Co metal synergistically enhances electron back-donation effects. The RhCo₃ cluster exhibits different adsorption and kinetic behaviors across the triplet, quintet, and septet potential energy surfaces, among which the septet state shows the most favorable catalytic performance. Energy span (δE) model analysis further indicates that, at low and moderate temperatures, the N₂ adsorption is the key factor governing the catalytic activity of RhCo₃. At 298 K, the reaction displays a δE value of 2.03 eV, and the catalytic activity increases with the temperature. However, at higher temperatures, the NH₃ desorption becomes the rate-determining process, shifting the turnover-determining transition state beyond the turnover-determining intermediate, and the energy span increases to 2.95 eV. These findings elucidate the temperature-dependent catalytic mechanism of RhCo₃ and provide theoretical insights for the rational design of efficient bimetallic catalysts for ammonia synthesis.

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1. Introduction

Ammonia (NH₃) is a crucial raw material used in the fertilizer, pharmaceutical, and chemical industries.^{1,2} Owing to its high hydrogen content, NH₃ holds promise as a carbon-free energy storage fuel and is increasingly becoming a valuable resource in future energy systems.^{3–5} The low electron affinity (–1.8 eV), nonpolar nature, and high ionization energy (15.58 eV) of the nitrogen molecule make its activation challenging.⁶ Therefore, this process requires high-temperature and high-pressure conditions, leading to significant CO₂ emissions and substantial energy consumption. In fact, the energy-intensive nature of ammonia production (primarily *via* the Haber–Bosch process) stimulates further research into more sustainable and efficient alternatives. The electron-donating ability of the catalyst has been shown to be crucial for the activation of N₂. The development of a highly efficient catalyst for the nitrogen cycle could result in substantial energy savings and reduced consumption in the ammonia synthesis industry.⁷

Multimetallic cluster catalysts have the potential to achieve high catalytic activity,^{8–11} thanks to the enhanced synergy

among multiple metals. The metal centers of these catalysts increase the electron-donating ability to the π^* orbitals of N₂ antibonds,¹² thereby weakening the N≡N bond and facilitating the ammonia synthesis.^{13,14} This synergy results in a superior catalytic performance compared to single-metal catalysts;¹⁵ transition metals such as rhodium (Rh) and cobalt (Co), along with Fe and Ru, exhibit high activity in activating N₂, as demonstrated by their volcano-type curves.^{16,17} However, Rh-based catalysts have several significant drawbacks, including high cost and poor stability, which limit their broad application.

The partially filled d orbitals of transition metals such as Rh and Co provide unique electronic modulation abilities, making them highly effective catalytic centers for small-molecule activation.¹⁸ Recent theoretical and experimental studies have shown that spin-related properties can influence the thermodynamics and kinetics of catalytic reactions.^{19–21} While Rh and Co are located in different rows of the periodic table, they both belong to group VIII B, and their similar d-orbital electronic structures may induce d–d orbital interactions between them, enhancing their electron exchange and transfer properties with adsorbed species or intermediates.²² Furthermore, electron spin effects are crucial in regulating catalytic reactions; the incorporation of Co atoms into Rh active sites can create a bimetallic synergy, enhancing the spin effects.^{23–25} Thus, the development of RhCo₃ catalysts, harnessing d–d orbital coupling to enhance

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spin interactions, represents a promising strategy for activating N₂-based reactions. Nevertheless, the mechanistic interplay between synergistic and spin effects in bimetallic transition metal catalysts remains poorly understood. In this work, we employ quantum chemical calculations and a d-d orbital coupling strategy to explore the cooperative spin effects of isolated RhCo₃ clusters in the N₂ to NH₃ conversion, focusing on the modulation of electron spins. Moreover, we apply the energy span model proposed by Kozuch to calculate the turnover frequency (TOF), establishing a framework for optimizing the performance of bimetallic catalysts.

2. Computational methods and theoretical background

2.1 Molecular geometry optimization

All calculations were carried out using the Gaussian 16 program.²⁶ Density functional theory (DFT)²⁷ calculations with the B3LYP functional²⁸ were performed using the SDD pseudopotential basis set²⁹ for the Co and Rh transition metals, while the 6-311++G(2df, 2pd) basis set was employed for the H and N atoms. The molecular geometric structures of all species within the reaction system, including reactants, intermediates, transition states, and products, were optimized and confirmed by frequency analysis. The thermodynamic properties, zero-point energies, and vibrational frequencies of the stable structures were then determined. The energies of reactants, intermediates, and products were confirmed as local minima (with no imaginary frequencies), while transition states exhibited unique imaginary vibrational frequencies. The conversion from transition states to reactants and products was analyzed using the intrinsic reaction coordinate (IRC) method to confirm the structural integrity. Wavefunction analysis of the reaction system was conducted using the Multiwfn package.³⁰ The bond dissociation process was studied through electrostatic potential (ESP) theory,³¹ Mayer bond order analysis,³² and the electron localization function (ELF) method.^{33,34}

2.2 Calculation of adsorption and dissociation energies

The adsorption and dissociation energies ($E_{\text{ads}}(\text{X})$ and $E_{\text{diss}}(\text{Y})$, respectively) reflect the interaction strength of adsorbates on catalysts and the energy required for their dissociation, respectively. They were calculated using the following equations:³⁵

$$E_{\text{ads}}(\text{X}) = E_{\text{X/catalyst}} - (E_{\text{catalyst}} + E_{\text{X}}) \quad (1)$$

$$E_{\text{diss}}(\text{Y}) = E_{\text{catalyst}} + E_{\text{Y}} - E_{\text{im}} \quad (2)$$

where $E_{\text{X/catalyst}}$ and E_{catalyst} are the energy of the adsorption complex and isolated catalyst molecule, respectively; moreover, E_{X} and E_{Y} represent the energies of the gas-phase adsorbed reactant X and dissociated product Y molecules, respectively. E_{im} is the energy of the pre-dissociation reaction intermediate (e.g., an intact ammonia molecule adsorbed on the catalyst surface).

2.3 Calculation of TOF

The TOF reflects the activity of a catalyst, and was calculated based on the energy span model, as described by the following equations:^{36,37}

$$\text{TOF} \approx \frac{k_{\text{B}}T}{h} e^{(-\delta E/RT)} \quad (3)$$

$$\delta E = T_{\text{TDS}} - I_{\text{TDI}} + \delta G_{ij} \quad \delta G_{ij} \begin{cases} 0 & T_{\text{TDS}} \text{ after } I_{\text{TDI}} \\ \Delta G_{\text{r}} & T_{\text{TDS}} \text{ before } I_{\text{TDI}} \end{cases} \quad (4)$$

where T_{TDS} and I_{TDI} represent the Gibbs free energies of the TDS (TOF-determining transition state) and the TDI (TOF-determining intermediate), respectively, while δG_{ij} represents the total Gibbs free energy change of the reaction.

The degree of TOF control (X_{TOF}) is a quantitative parameter used to evaluate the extent to which each reaction intermediate or transition state influences the overall reaction rate. X_{TOF} was calculated using the following equation:

$$X_{\text{TOF}, T_i} = \frac{\sum_j e^{(T_i - I_j - \delta G_{ij})/RT}}{\sum_{ij} e^{(T_i - I_j - \delta G_{ij})/RT}}, \quad X_{\text{TOF}, I_{i=j}} = \frac{\sum_i e^{(T_i - I_j - \delta G_{ij})/RT}}{\sum_{ij} e^{(T_i - I_j - \delta G_{ij})/RT}} \quad (5)$$

3. Results and discussion

3.1 Analysis of d-d orbital coupling mechanism between Rh and Co in RhCo₃

As shown in Fig. 1(A), 5s electrons have a lower energy than 4d electrons, owing to shielding and penetration effects. The electron configuration of Rh is [Kr]4d⁸5s¹. According to Hund's rule, electrons fill equivalent orbitals with parallel spins, and electrons located within the same orbital have opposite spins (Pauli exclusion principle). Hence, the electrons will preferentially fill the 5s orbital and continue to fill the 4d orbitals (d_{xy} , d_{yz} , d_{xz} , $d_{x^2-y^2}$, d_{z^2}) to form the quartet (⁴Rh, ground) state of Rh. Excitation of a single electron from the 5s to the 4d orbital changes the electronic configuration to 4d⁹5s⁰, resulting in a doublet state (²Rh).

The electron configuration of Co is [Ar]3d⁷4s². Owing to the greater penetration effect of the 4s orbital, its energy is lower than that of the 3d orbital, causing electrons to occupy the 4s orbital first, and leading to a low-activity quartet ground state (⁴Co) with three unpaired electrons in the 3d orbital. Low-energy excitation of the 4s electrons of Co results in the formation of a quartet 3d⁸4s¹ configuration with a single unpaired electron in the 4s orbital, making this configuration more chemically reactive compared to the 3d⁷4s² configuration with two paired electrons in the 4s orbital. Co atoms in this state exhibit higher reactivity. At significantly higher excitation energies, all 4s electrons may move to the 3d orbitals, leading to a 3d⁹4s⁰ electronic configuration, corresponding to a doublet state (²Co) with one unpaired electron. This highly excited state, which is challenging to achieve in isolated atoms and unstable, may transiently form in RhCo₃ complexes.³⁸ As shown in Fig. 1(B), the Rh 4d and Co 3d orbitals can generate a strong d-



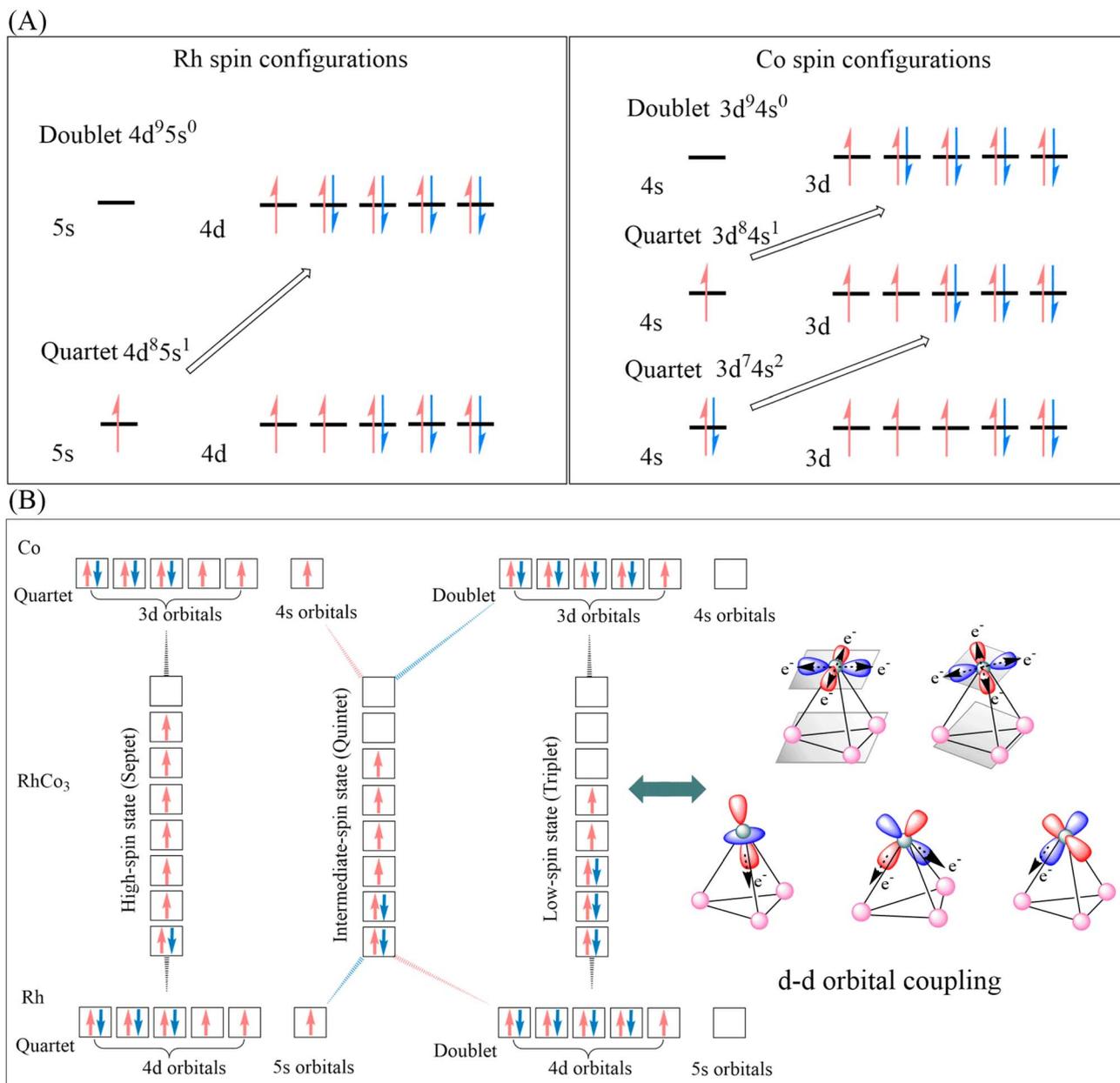


Fig. 1 (A) Spin configurations of Rh/Co; (B) d–d orbital coupling mechanism of RhCo₃.

d coupling. The strong spin–orbit coupling (SOC) causes the originally fivefold-degenerate d orbitals of Rh and Co to split into separate energy levels, forming two possible electronic configurations. The first configuration involves a face-to-face electron transfer between the Rh d orbitals and Co, where Rh contributes to form RhCo₃ through the $d_{x^2-y^2}$ and d_{xy} orbitals, corresponding to the higher-energy doubly degenerate orbitals. The other configuration involves the coupling of Rh d_{z^2} , d_{xz} , and d_{yz} orbitals with the Co d orbitals *via* vertex-to-vertex electron transfer, forming the lower-energy triply degenerate orbitals. Different filling patterns of d orbitals and combinations of spin orientations result in distinct spin multiplicity states of RhCo₃.

3.1.1 ³[RhCo₃]. As shown in Fig. 2(A), Co exhibits a mixture of high- and low-spin states. Specifically, the Co2 atomic

orbitals contain 1.96 α -spin electrons, while Co3 and Co4 contain β -spin electrons with populations of -0.22 and -0.26 , respectively. The Rh atoms possess approximately 0.52 α -spin electrons, which can interact with the Co ligands to form weak spin-paired bonds with Co3 and Co4.

3.1.2 ⁵[RhCo₃]. Co2 and Co4 atoms contribute 1.87 and 1.91 α electrons, respectively, while the Co3 atoms possess -0.31 β electrons. The asymmetric distribution of α and β electrons in the quintet RhCo₃ state has two effects. The high-spin Co2/Co4 atoms facilitate H transfer through the excess α electrons, whereas the β electrons of Co3 can form weak covalent interactions with the d orbitals of Rh to modulate the stability of reaction intermediates.



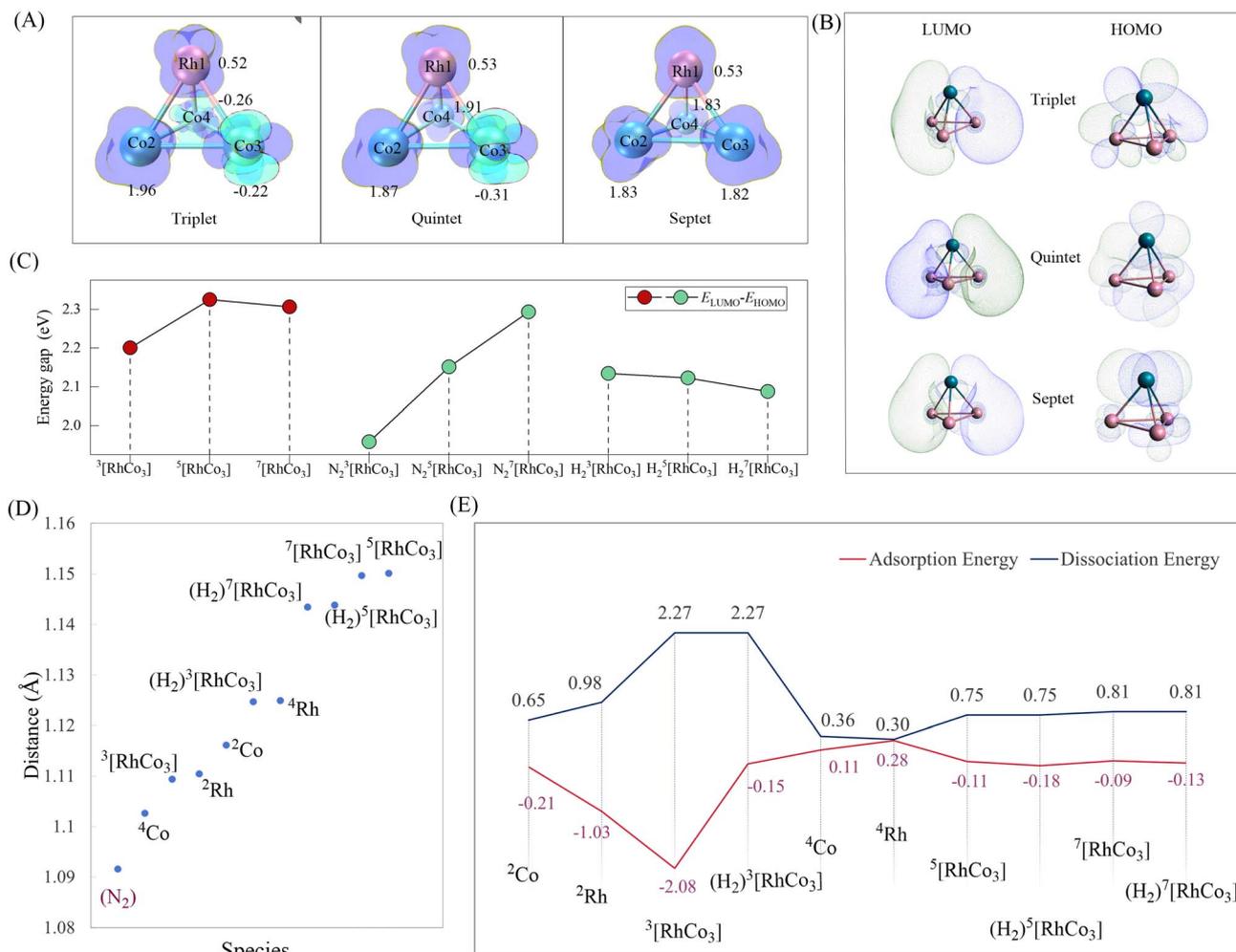


Fig. 2 (A) Electron spin density maps of RhCo₃, the atomic symbols (Rh1, Co2, Co3, Co4) are atomic labels within the molecular structure; (B) illustration of frontier molecular orbitals; (C) energy gap (ΔE) diagram of frontier molecular orbitals for various catalysts; (D) N≡N bond lengths of N₂ adsorbed on various catalysts; (E) N₂ adsorption and NH₃ dissociation energies on various catalysts.

3.1.3 ⁷[RhCo₃]. Rh and Co are coordinated in a quartet state with spin populations of 0.53, 1.83, 1.82, and 1.83 for Rh, Co2, Co3, and Co4, respectively. The d orbitals of Rh and Co contribute 0.72 and 6.28 α electrons, with the spin populations of s and p orbitals being -0.60 and -0.41. The total spin multiplicity is predominantly influenced by the strong spin polarization of the d orbitals. The disappearance of the β electrons suggests that the α electrons of ⁷[RhCo₃] are strongly delocalized. Although the single-electron densities at the Rh center in the triplet, quintet, and septet states of RhCo₃ are similar, the unpaired electron of the Co atom exhibits a distinct behavior, highlighting the different properties of the RhCo₃ bimetallic catalyst across different metal sites and spin states, particularly evident from the varying electron density at the Co atoms.

The frontier molecular orbitals (FMOs) of RhCo₃ are formed through the contribution of d electrons from Rh and Co (Fig. 2(B)). DFT calculations reveal that the adsorbed N₂/H₂ molecules induce additional splitting of Rh 4d and Co 3d orbitals in RhCo₃, resulting in non-degenerate highest occupied molecular orbital (HOMO) and lowest unoccupied molecular

orbital (LUMO) levels. According to FMO theory, a smaller energy gap between frontier orbitals (HOMO–LUMO gap) results in a higher reactivity. The energy gaps between the FMOs after adsorbing N₂ and H₂ are shown in Fig. 2(C). The ³[RhCo₃] cluster exhibits the smallest energy gap (2.20 eV \rightarrow 1.96 eV) upon N₂ adsorption, while ⁷[RhCo₃] shows the narrowest gap (2.31 eV \rightarrow 2.09 eV) after H₂ adsorption. The reduced post-adsorption gaps observed in both cases indicate enhanced catalytic activities, with ³[RhCo₃] and ⁷[RhCo₃] showing the highest N₂ and H₂ activation abilities, respectively.

3.2 Adsorption properties of RhCo₃ and analysis of electrostatic potential interactions

According to the Sabatier principle,^{39–41} an effective catalyst should exhibit an optimal binding strength, balancing the adsorption of reactants and the desorption of products, with the reaction rate peaking at a moderate adsorption energy. The bond lengths of chemically adsorbed N and H species on the surfaces of Co, Rh, and RhCo₃ are compared in Fig. 2(D). Upon adsorption, electrons from the d orbitals of the Co, Rh, and RhCo₃



metals migrate to the antibonding orbitals of nitrogen, leading to an increase in the N≡N bond length. The $^5[\text{RhCo}_3]$ and $^7[\text{RhCo}_3]$ catalysts exhibit the highest activity, resulting in an increase in the N≡N bond length from 1.09 to 1.15 Å upon N_2 adsorption, whereas that of $(\text{H}_2)^5[\text{RhCo}_3]$ and $(\text{H}_2)^7[\text{RhCo}_3]$ increases to 1.14 Å. RhCo_3 exhibits distinct active site characteristics near Rh and Co sites, as evidenced by their different catalytic behaviors. The d orbitals of Rh are partially filled and remain unsaturated within the complex, facilitating the spontaneous adsorption of N_2 and H_2 on the active Rh sites of RhCo_3 . $^3[\text{RhCo}_3]$ displays the lowest N_2 adsorption energy (-2.08 eV), but at the same time makes the NH_3 desorption unfavorable (dissociation energy 2.27 eV) (Fig. 2(E)). $^7[\text{RhCo}_3]$ exhibits thermodynamic benefits (SI Tables S1–S3) in addition to a moderate N_2 adsorption energy (-0.09 eV); following H_2 adsorption, the N_2 adsorption energy on $(\text{H}_2)^7[\text{RhCo}_3]$ decreases to -0.13 eV, indicating a good resistance to hydrogen poisoning. The NH_3 desorption is a crucial step in the ammonia synthesis; the calculated NH_3 desorption barrier from $^7[\text{RhCo}_3]$ (0.81 eV) is lower than that of catalysts such as FeMo (1.00 eV) and Fe_3Co (0.90 eV), and close to that of the highly active RuMo (0.77 eV).⁴² This indicates that $^7[\text{RhCo}_3]$ exhibits favorable performance in facilitating NH_3 desorption.

The extrema of the electrostatic potential for each catalyst are shown in Fig. 3(A). The N site displays a negative potential ($V_{s,\text{min}} = -0.35$ eV) on its terminal side, attributed to lone-pair electrons, while the bridging side shows a $V_{s,\text{max}}$ value of 0.31 eV. The minimum potential energy for H_2 bridging adsorption ($V_{s,\text{min}}$) is -0.12 eV. The RhCo_3 cluster shows an increased separation between the positive and negative electrostatic potential regions, owing to the electron cloud density shifting toward the more electronegative Rh. This separation becomes more pronounced with increasing single-electron density. The $^7[\text{RhCo}_3]$ cluster exhibits a higher spin polarization (with more unpaired electrons), generating an ESP gradient of 2.31 eV between its positive and negative extrema, which significantly increases the reactivity of its polarized regions, as shown in Fig. 3(B). Both N_2 and H_2 adsorb onto the Rh active center in a bridging configuration. The size of the overlapping electrostatic potential regions between RhCo_3 and N_2 in the triplet, quintet, and septet states exhibits a gradual increase (1.91 → 2.02 → 2.14 Å), while those of the regions between H_2 and RhCo_3 show a gradual decrease (2.34 → 2.17 → 2.11 Å). This trend implies that a higher single-electron spin density results in stronger N_2 and weaker H_2 adsorptions. The

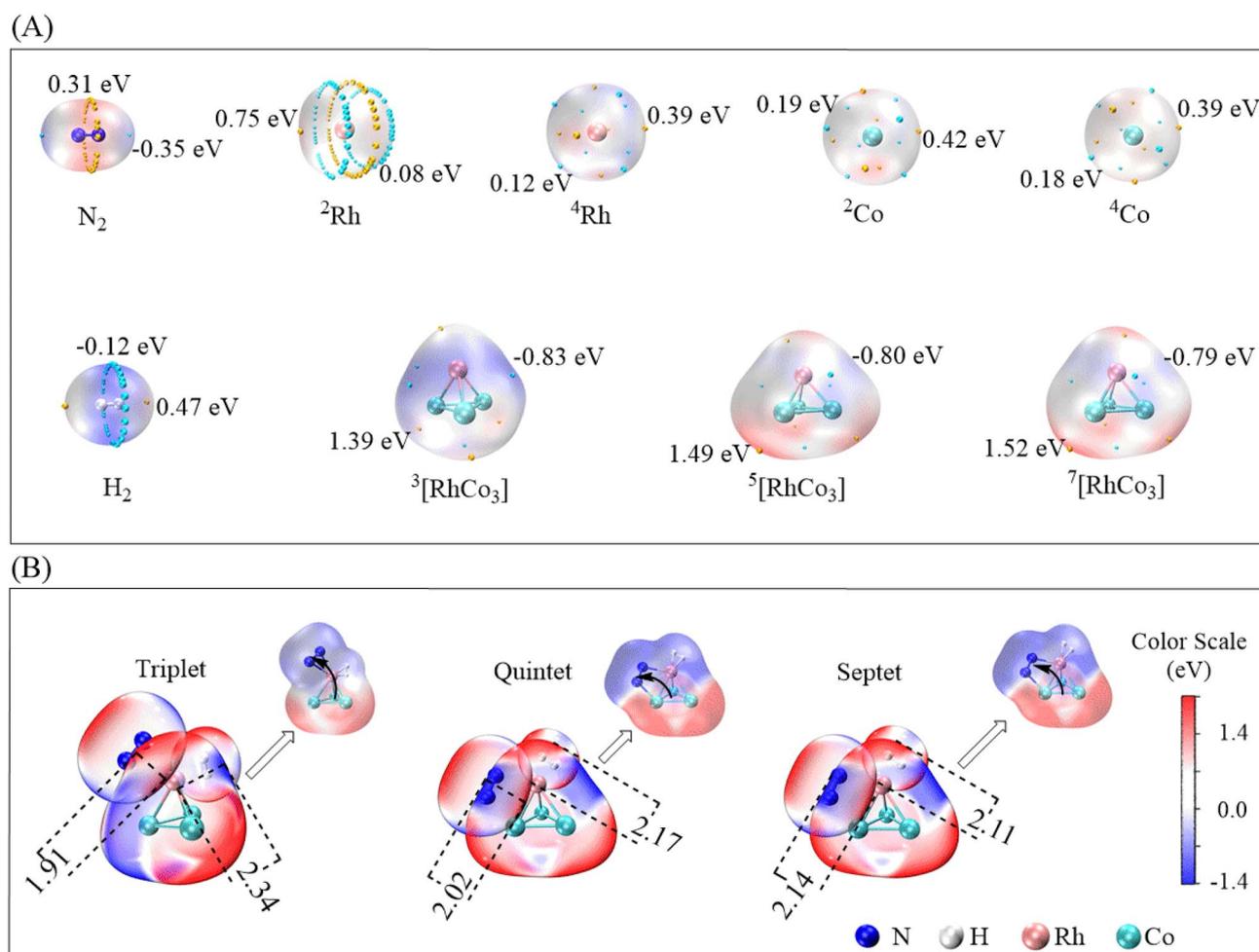


Fig. 3 (A) Illustration of electrostatic potential extrema; (B) electrostatic potential penetration of co-adsorbed N_2 and H_2 on RhCo_3 cluster surface (units: Å).



electronegativity of N causes a shift of the electron density of the Rh-Co cluster toward the outer side of the Rh center in the N direction, leading to its intensification near the -NN moiety. The elongation of the N-N bond and the increased electronegativity of N suggest a weakening of the N≡N bond. This shows that the multiple-spin single electrons of RhCo₃ can enhance the adsorption ability of Rh within the metal cluster, activating the N-N bond.

3.3 Molecular structure optimization and reaction PES calculation

The competitive adsorption of H₂ is thermodynamically favored over that of N₂, leading to the sequential adsorption of H₂ and N₂ at the Rh site. The reaction potential energy surfaces (PESs)

for the triplet, quintet, and septet states are shown in Fig. 4(A), and the corresponding energies and geometries are provided in the SI (Fig. S1-S3).

In the uncatalyzed pathway, the transfer of the first H atom is the rate-determining step for the N₂ reaction with H₂, with an energy barrier of 5.43 eV.⁴³ This is the step with the highest energy barrier in the sequential insertion process of three H₂ molecules. The catalytic H transfer (TS1-2) reaction on isolated RhCo₃ exhibits the highest energy barrier (3.41 eV) on the triplet PES, with lower barriers of 1.92 and 1.81 eV observed on the quintet and septet surfaces, respectively. The TS1-2 vibrational frequencies corresponding to different spin states exhibit the typical characteristics of energy barriers for H transfer reactions (*i.e.*, high and thin barriers). As the number of unpaired spin

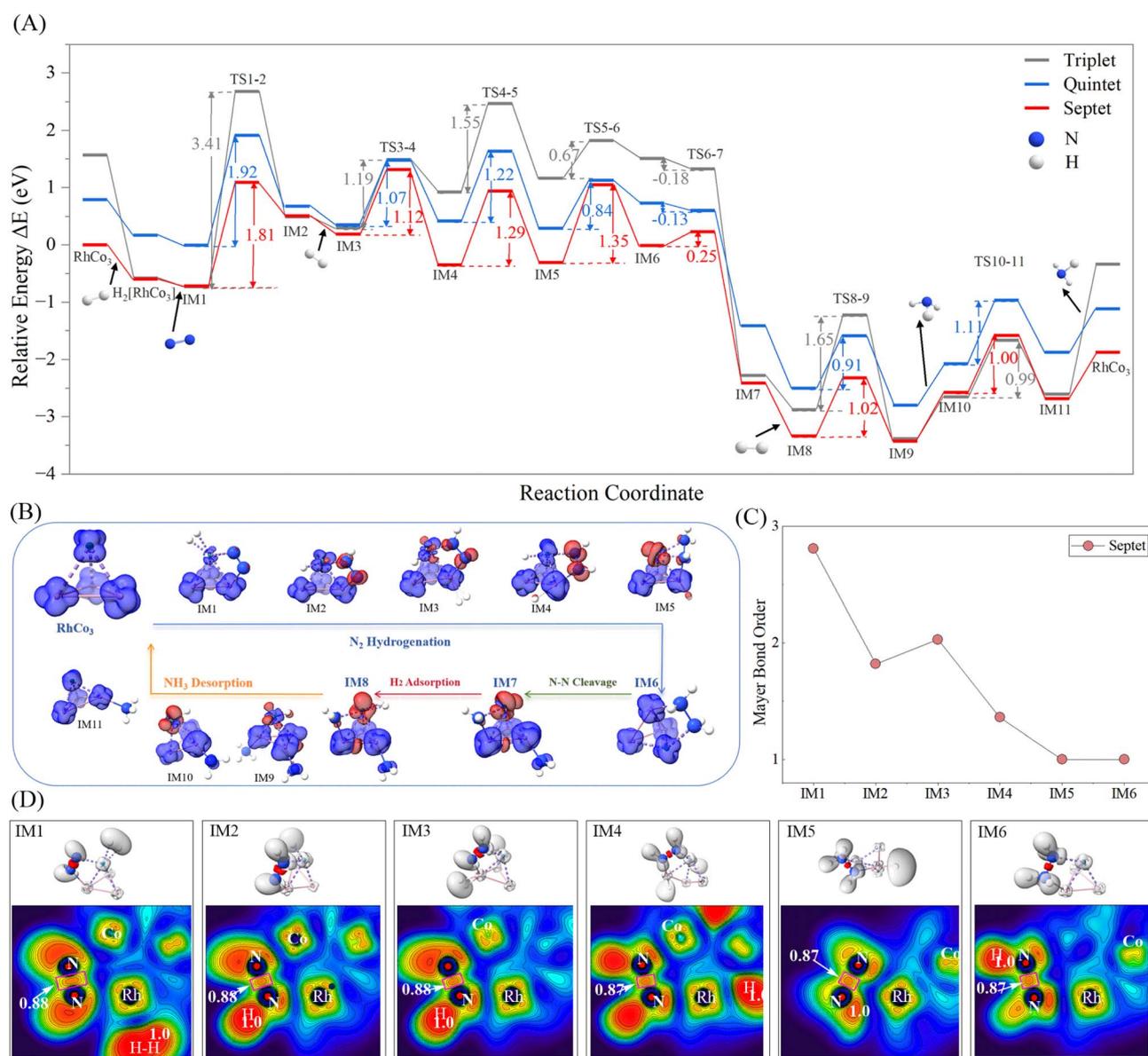


Fig. 4 (A) PESs of N₂ and H₂ reactions catalyzed by triplet/quintet/septet RhCo₃; (B) reaction mechanism of N₂ activation by ⁷[RhCo₃] and illustration of spin density changes; (C) N-N bond order of partial intermediate species formed on septet RhCo₃; (D) electronic localization maps of partial intermediates formed on septet RhCo₃.



electrons increases, the repulsive interaction between these electrons causes an increase in the Rh–N bond distance, strengthening the exchange interaction between the electrons. This leads to a decrease in the force constants, with the vibrational frequencies corresponding to the H transfer decreasing in the order triplet (-918.61 cm^{-1}) > quintet (-839.38 cm^{-1}) > septet (-783.93 cm^{-1}) (Tables S4–S6). The N hydrogenation reaction generally follows either the alternating hydrogenation or the distal associative hydrogenation pathway. The second H adsorption is the key step in the hydrogenation mechanism. Owing to the presence of the pre-adsorbed H₂, the Rh site in IM2 exhibits significant steric hindrance. The Co2 site provides additional adsorption space, and the induced positive charge on the NNH moiety (IM2 0.06 eV, Table 1) facilitates the H transfer from Co2 to N, thereby determining the alternating hydrogenation reaction pathway. In this process, the PESs of the three spin states for the H transfer process exhibit similar shapes. The two N atoms undergo successive hydrogenation steps, leading to gradual elongation of the N–N bond and cleavage of the two π bonds between the N atoms. After four H transfers, the RhH₂N–NH₂Co2 (IM6) species is formed. The transition from IM6 to IM7 involves the cleavage of the N–N σ bond. This is a barrierless reaction in the triplet state ($E_{\text{TS6-7}} < E_{\text{IM6}}$), whereas it requires overcoming a barrier of 0.25 eV in the septet state. After the N–N bond cleavage, the system energy significantly decreases, with the NH₃ dissociation occurring at the Co3 and Co2 sites. The final products are RhCo₃ and NH₃, and the overall catalytic cycle is exothermic ($\Delta E = -1.9\text{ eV}$).

The DFT-calculated PESs corresponding to different spin states reveal that the triplet exhibits a more stable adsorption intermediate structure. This stability is due to Rh and Co forming a strong ligand field environment with low spin states in the triplet configuration, with Rh serving as the electron donor (Table 1) and carrying a positive charge, while Co bears a partial negative charge. During H transfer and NH₃ dissociation, the Rh–N and Co–NH₃ bonds exhibit shorter lengths due to their high stability under strong ligand field effects, making them highly thermodynamically stable, but kinetically unfavorable. As a result, transition states on the triplet-state PES generally feature higher energy barriers, requiring more energy for NH₃ dissociation. In the septet state, both Rh

and Co adopt a high-spin coordination, promoting the formation of adsorption intermediates with comparatively lower energies. Fig. 4(B) illustrates the changes in the electron spin density; the adsorption of H₂ induces electron accumulation (red electron cloud) on the Rh surface (IM2) and transfer to the N antibonding orbital, which reduces the N–N bond order and weakens the N≡N bond. This, in turn, leads to the progressive weakening of the π and σ bonds of N₂. The reaction mechanism can be divided into four distinct stages: N₂ hydrogenation → N–N bond cleavage → H₂ adsorption → NH₃ dissociation. As shown in Fig. S4, the variations in metal–metal bond lengths along the reaction coordinate indicate that the RhCo₃ cluster remains generally stable throughout the catalytic process. The Rh–Co bond lengths exhibit only slight fluctuations, suggesting that the metallic framework of the catalyst possesses high structural strength. A temporary distortion of the Co2–Co4 bond occurs for the IM4 intermediate, likely originating from local structural rearrangements associated with H migration and N₂ activation. However, the system subsequently recovers its original configuration, demonstrating the excellent structural flexibility and catalytic stability of the RhCo₃ catalyst.

As the H transfer proceeds, the N–N bond order gradually decreases from 2.81 (co-adsorption state) to 1.82, 2.03, 1.36, and eventually to 1, until complete bond cleavage occurs for the IM7 intermediate, as clearly shown in Fig. 4(C) for the catalytic reaction in the septet state. The electron localization analysis shown in Fig. 4(D) reveals localized electron clouds (red) with covalent bond characteristics between the two N atoms of the N₂ molecule; moreover, localized electron clouds (gray) are also present between H atoms during the H₂ adsorption on IM1 and IM2. The ELF value of the N–N bond stabilizes at 0.88, while the electron-localized region of the N–H bond progressively expands as the reaction proceeds. Concurrently, the red-colored covalent character region of the N–N bond gradually shrinks, corresponding to the sequential reduction in bond order during the H-atom transfer.

3.4 Bimetallic synergistic effect of Rh and Co

According to the charge distribution in the septet reaction system, Rh primarily acts as the electron-donating center, with a positive electron cloud density (Table 1).

In RhCo₃, Rh carries a positive charge of 0.37, which results in a strong adsorption of H₂ molecules. In the co-adsorbed IM1, the electron densities of Rh1, Co2, Co3, and Co4 are all reduced, and the Co atoms may receive electrons back-donated from the N atom *via* the Rh center, thereby influencing the electron cloud densities of both Rh and Co atoms. Mulliken charge analysis reveals that the Co2 atom in the IM2 intermediate carries a net positive charge of +0.06, demonstrating its H₂ adsorption ability. As the positive charge on Rh increases and its electron cloud density decreases (in IM1, IM2, IM4, and IM8), the electron cloud density of Co exhibits a simultaneously decrease. In IM9 and IM11, after the formation of the NH₃ molecule, Co donates electrons back to N, leading to an increase in the electron cloud density of Co.

Table 1 Atomic charge distribution of intermediates on septet PES

Atomic charge	Rh1	Co2	Co3	Co4
RhCo ₃	0.37	−0.13	−0.12	−0.12
IM1	0.64	−0.10	−0.02	−0.06
IM2	0.79	0.06	−0.10	−0.12
IM3	0.85	−0.21	−0.05	−0.14
IM4	1.24	−0.08	−0.03	−0.11
IM5	0.87	−0.07	−0.23	−0.11
IM6	0.24	−0.09	−0.20	−0.22
IM7	0.77	−0.44	0.16	0.03
IM8	1.12	−0.43	0.20	0.06
IM9	0.98	−0.31	−0.42	0.09
IM10	0.76	−0.29	0.02	0.08
IM11	0.49	−0.55	−0.09	−0.09



Table 2 Spin population of intermediates on septet PES

Spin population	Rh1	Co2	Co3	Co4
RhCo3	0.53	1.83	1.82	1.83
IM1	0.03	1.99	2.04	2.04
IM2	0.34	1.96	2.04	2.10
IM3	0.07	1.79	2.10	2.32
IM4	0.32	2.38	1.91	2.17
IM5	-0.65	2.56	2.13	2.02
IM6	0.54	1.77	1.87	1.82
IM7	-0.76	2.48	2.22	2.00
IM8	-0.59	2.51	2.10	1.91
IM9	-0.22	2.33	2.16	1.78
IM10	-0.61	2.58	1.98	1.99
IM11	0.37	2.01	1.84	1.84

The results indicate that the Co and Rh atoms can synergistically influence the catalytic reactions by dynamically adjusting the electron cloud around the Rh atoms to meet the reaction requirements. Table 2 shows that the total spin population of all intermediates in the septet state remains around 6, with the spin population of Rh ranging from -0.76 to 0.53. In some of the intermediates (IM5, IM7, IM8, IM10), the transfer of d electrons from Rh to N leads to a negative spin population for the Rh center, while that of the Co2 atoms significantly increases to above 2.4 (2.56, 2.48, 2.51, 2.58) to maintain the overall spin stability. During the H-to-N transfer process, the Rh atom in RhCo₃ predominantly acts as the reactant activation site by engaging with the reactant molecules and activating the N≡N bond. The Co atom supports this role by dynamically adjusting the electron cloud environment around Rh through electron transfer and spin modulation, and potentially aiding in maintaining the structural stability of the active center of the catalyst.

3.5 Analysis of reaction pathways and catalytic TOF

Fig. 5 shows the relative Gibbs free energy diagrams (298 K) for the RhCo₃-catalyzed N₂ + 3H₂ → 2NH₃ reaction across the triplet, quintet, and septet states. The total spin quantum

number difference $\Delta m_s = \pm 1$ for the triplet and quintet states allows potential energy surface crossings, with multiple crossing points (CPs) distributed before the H₂ adsorption and H transfer barriers. However, the energy gaps of the adsorbates and transition states vary significantly (ranging from 0.99 to 2.64 eV), requiring the overcoming of high energy barriers, and the energies of all states remain higher than those observed on the septet PES. For the triplet-septet transition, $\Delta m_s = \pm 2$, which is forbidden, implying minimal PES mixing at crossing points. The Gibbs free energy barriers for consecutive hydrogenations over the septet state ⁷[RhCo₃] are 1.81, 1.13, and 1.02 eV. Compared with the Nb-based bimetallic system Nb₂Rh (1.70, 1.74, and 1.80 eV), ⁷[RhCo₃] exhibits lower energy barriers for the sequential hydrogenation steps, showing a notable kinetic advantage in the second and third hydrogenation processes.⁴⁴

Theoretical calculations with the energy span model were performed to determine the TOF values and rate-determining states of the RhCo₃-catalyzed N₂ and H₂ cyclic reactions. The calculations reveal that the N₂ activation reaction over septet RhCo₃ exhibits high thermal sensitivity within the temperature range of 298–478 K. The turnover-determining intermediate (TDI) is identified as IM1, while the turnover-determining transition state (TDTS) corresponds to TS3-4, defining the rate-determining region between these two states. The adsorption of N₂ and the subsequent cleavage of the N–N σ bond are the key steps influencing the overall reaction rate. The energy span (δE) is calculated to be 2.03 eV, and the TOF of RhCo₃ increases exponentially from 2.98×10^{-22} to $2.53 \times 10^{-9} \text{ s}^{-1}$, following a typical Arrhenius behavior (Fig. 6). When the temperature exceeds 478 K, the relative free energy values of the intermediates changes: the TDTS remains positioned at TS3-4, while the TDI shifts to IM9, which is located beyond the TDTS. Hence, the energy span increases to $\delta E = T_{\text{TDTS}} - I_{\text{TDI}} + \delta G_{ij} = 2.95 \text{ eV}$, which may lead to a decline in the TOF. This result suggests that, at higher temperatures, the influence of the N–N bond dissociation on the RhCo₃ activity is reduced, and the desorption step becomes the dominant process.

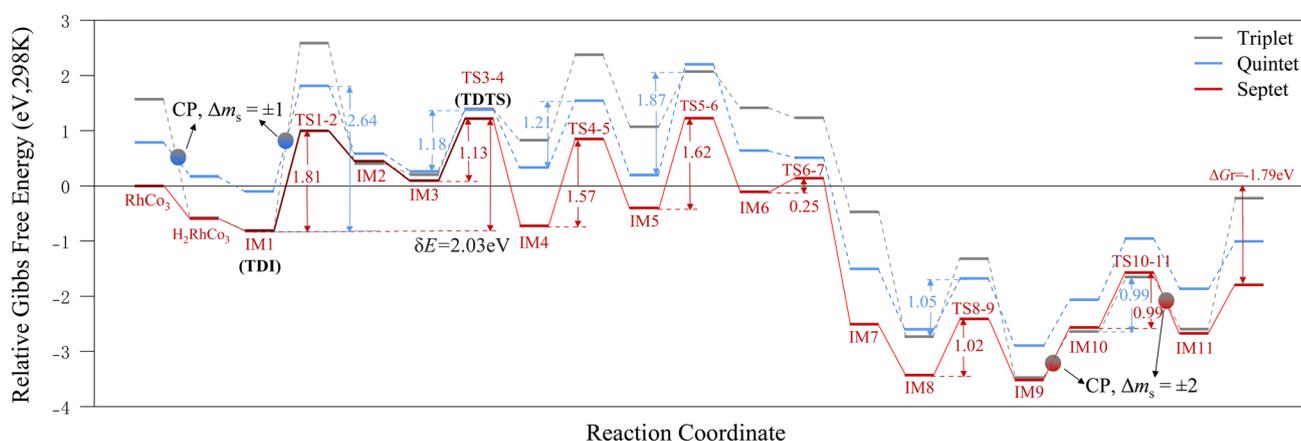


Fig. 5 Relative Gibbs free energy diagrams for RhCo₃-catalyzed N₂ + 3H₂ → 2NH₃ reaction at 298 K.



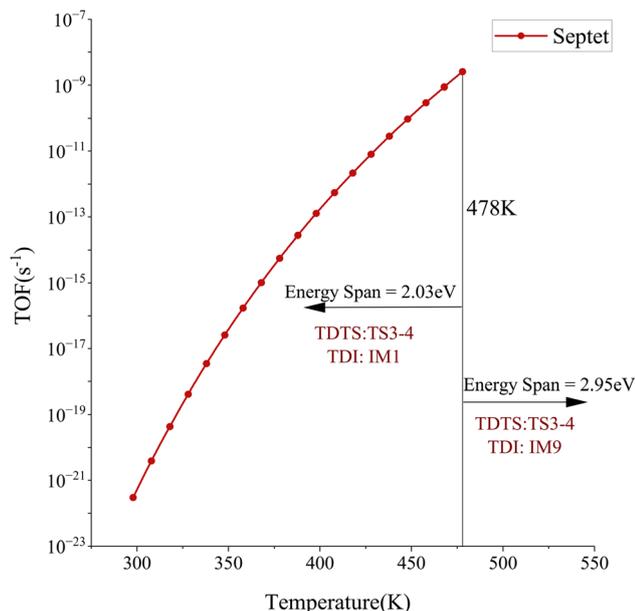


Fig. 6 The relationship between the TOF and temperature for the ${}^7[\text{RhCo}_3]$ -catalyzed N_2 reaction for ammonia synthesis.

4. Conclusions

In this study, we have investigated the mechanism of the $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ reaction promoted by the d–d coupling-based bimetallic catalyst RhCo_3 , focusing on the synergistic and spin effects between Rh and Co. The results show that the unsaturated d orbitals of Rh and Co atoms in RhCo_3 , combined with the significant electronegativity difference, induce a marked d–d orbital coupling. This interaction generates a favorable bimetallic synergy and induces spin effects, which collectively modulate the N_2 activation ability. Septet RhCo_3 exhibits an H_2 adsorption energy of -0.59 eV, a subsequent N_2 adsorption energy of -0.13 eV, as well as NH_3 dissociation energies of 0.85 and 0.81 eV, indicating moderate adsorption and dissociation energies with low energy barriers for H transfer. In the H transfer process to N, the Rh atom acts as electron regulation center and reactant activation site, while the Co metal contributes to enhance the electron back-donation effect. This synergistic effect is primarily attributed to the establishment of a dynamic electron-transfer mechanism between Rh and Co within the polymetallic centers of the RhCo_3 cluster. Between 298 and 478 K, the septet RhCo_3 catalyst exhibits a significant increase in TOF, peaking at 478 K. The d–d coupling-based bimetallic catalyst RhCo_3 effectively promotes N_2 activation through bimetallic synergy and spin effects. This catalyst achieves an optimal balance between the adsorption energies of reactants and key intermediates, the product dissociation energies, and the H transfer barriers, enabling an efficient NH_3 synthesis. The spin effects of Rh and Co play a critical role in the catalytic process. In conclusion, this study provides theoretical support for the design of multimetallic catalysts with properties tunable *via* spin effects.

Author contributions

Jingjing Wu carried out the investigation and wrote the manuscript. Haixiong Shi contributed to supporting roles in the literature survey. Yongcheng Wang developed the methodology.

Conflicts of interest

There are no conflicts to declare.

Data availability

The data supporting the findings of this study are available within the article and its supplementary information (SI). Supplementary information: optimized geometries (Cartesian coordinates), energy values, and transition-state vibrational frequencies. See DOI: <https://doi.org/10.1039/d5ra06945a>.

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References

- S. Voelker, N. Groll, M. Bachmann, L. Mueller, M. Neumann, T. Kossioris, P. Muthyala, B. Lehrheuer, M. Hofmeister, A. Vorholt, K. Schmitz, S. Pischinger, W. Leitner and A. Bardow, Towards Carbon-Neutral and Clean Propulsion in Heavy-Duty Transportation with Hydroformylated Fischer–Tropsch Fuels, *Nat. Energy*, 2024, **9**(10), 1220–1229.
- H. Guo, A. R. P. Harrison, M. Gao, X. Zhang, Q. Chen, Z. Cui and B. Nie, Chemical Looping Based Low-Pressure Ammonia Synthesis, *Chem. Eng. J.*, 2024, **500**, 157321.
- M. T. Powders, B. A. Luqmani, M. Pidou, M. Zhu and E. J. McAdam, The Use of Ammonia Recovered from Wastewater as a Zero-Carbon Energy Vector to Decarbonise Heat, Power and Transport - A Review, *Water Res.*, 2025, **268**, 122649.
- R. Tort, A. Bagger, O. Westhead, Y. Kondo and A. Khobnya, Nitrogen Reduction to Ammonia: Drawing a Path Toward Energy Efficiency, *ECS*, 2024, **MA2024-01**(39), 2312.
- Y. Qi, W. Liu, S. Liu, W. Wang, Y. Peng and Z. Wang, A review on Ammonia-hydrogen fueled internal combustion Engines, *ETransportation*, 2023, **18**, 100288.
- H. P. Jia and E. A. Quadrelli, Mechanistic aspects of dinitrogen cleavage and hydrogenation to produce ammonia in catalysis and organometallic chemistry: relevance of metal hydride bonds and dihydrogen, *Chem. Soc. Rev.*, 2014, **43**, 547–564.
- C. M. Vanderham, M. M. Koper and D. H. Hettler, Challenges in reduction of dinitrogen by proton and electron transfer, *Chem. Soc. Rev.*, 2014, **43**(15), 5183–5191.



- 8 E. L. Muetterties, Metal Clusters in Catalysis III. - Clusters as Models for Chemisorption Processes and Heterogeneous Catalysis, *Bull. Soc. Chim. Belg.*, 2010, **84**(10), 959–986.
- 9 A. K. Raigar, Manju, K. Saini and A. Guleria, CuFe₂O₄-catalyzed one-pot synthesis of α -substituted 2-benzofuranmethamines *via* tandem A³coupling, 5-exo-dig cyclization, and 1,3-allylic rearrangement, *RSC Adv.*, 2025, **15**(39), 32108–32116.
- 10 E. L. Muetterties, E. L. Hoel and C. G. Salentine, Intramolecular rearrangements in boron clusters, *Inorg. Chem.*, 1975, **14**(4), 950–951.
- 11 E. L. Muetterties, Metal Clusters: Bridges between molecular and solid-state chemistry, *Chem. Eng. News*, 1982, **60**(35), 28–41.
- 12 T. Yamabe, K. Hori, T. Minato and K. Fukui, Theoretical study on the bonding nature of transition-metal complexes of molecular nitrogen, *Inorg. Chem.*, 1980, **19**, 2154.
- 13 M. D. Fryzuk and S. A. Johnson, The continuing story of dinitrogen action, *Coord. Chem. Rev.*, 2000, **200**, 379–409.
- 14 P. L. Holland and L. Patrick, Metal-dioxygen and metal-dinitrogen complexes: where are the electrons?, *Dalton Trans.*, 2010, **39**(23), 5415–5425.
- 15 P. Buchwalter, J. Rosé and P. Braunstein, Multimetallic catalysis based on heterometallic complexes and clusters, *Chem. Rev.*, 2014, **115**(1), 28–126.
- 16 A. A. Vojvodic, A. J. Medford, F. Studt, F. Abild-Pedersen, T. S. Khan, T. Bligaard and J. K. Nørskov, Exploring the limits: A low-pressure, low-temperature Haber-Bosch process, *Chem. Phys. Lett.*, 2014, **45**(25), 108–112.
- 17 A. Vojvodic and J. K. Nørskov, New design paradigm for heterogeneous catalysts, *Natl. Sci. Rev.*, 2015, **000**(002), 140–143.
- 18 R. C. Pang, P. F. Tian, H. L. Jiang, M. H. Zhu, X. Z. Su, Y. Wang, X. L. Yang, Y. H. Zhu, L. Song and C. Z. Li, Tracking structural evolution: operando regenerative CeOx/Bi interface structure for high-performance CO₂ electroreduction, *Natl. Sci. Rev.*, 2020, **8**, nwa187.
- 19 C. Biz, M. Fianchini and J. Gracia, Strongly correlated electrons in catalysis: focus on quantum exchange, *ACS Catal.*, 2021, **22**, 14249–14261.
- 20 J. Gracia, Spin dependent interactions catalyse the oxygen electrochemistry, *Phys. Chem. Chem. Phys.*, 2017, **31**, 20451–20456.
- 21 J. J. Wu, Y. C. Wang and J. X. Kang, A theoretical study for the synthesis of ammonia in reaction of N₂ and H₂ catalyzed by FeO, *Sci. Sin.: Chim.*, 2019, **49**(12), 1481–1490.
- 22 L. Zeng, Y. Chen, M. Sun, Q. Huang, K. Sun, J. Ma, J. Li, H. Tan, M. Li, Y. Pan, Y. Liu, M. Luo, B. Huang and S. Guo, Cooperative Rh-O5/Ni(Fe) site for efficient biomass upgrading coupled with H₂ production, *J. Am. Chem. Soc.*, 2023, **32**, 17577–17587.
- 23 X. C. You, Z. Y. Guo, Q. L. Jiang, J. K. Xia, S. W. Wang, X. H. Yang, Z. C. Zhuang, Y. F. Li, H. Xiang, H. Li and B. Yu, Magnetic-Field-Induced Spin Transition in Single-Atom Catalysts for Nitrate Electrolysis to Ammonia, *Nano Lett.*, 2025, **25**(21), 8704–8712.
- 24 K. Zhang, A. Cao, L. H. Wandall, V. Jerome, J. Kibsgaard, J. K. Nørskov and L. Chorkendorff, Spin-mediated promotion of Co catalysts for ammonia synthesis, *Science*, 2024, **383**(6689), 1357–1363.
- 25 H. Ren, G. F. Li, B. Zhu, X. D. Lv, L. S. Yao, X. L. Wang, Z. M. Su and W. Guan, How Does Iridium(III) Photocatalyst Regulate Nickel(II) Catalyst in Metallaphotoredox-Catalyzed C-S Cross-Coupling? Theoretical and Experimental Insights, *ACS Catal.*, 2019, **9**(5), 3858–3865.
- 26 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, G. A. Petersson, H. Nakatsuji, X. Li, M. Caricato, A. V. Marenich, J. Bloino, B. G. Janesko, R. Gomperts, B. Mennucci, H. P. Hratchian, J. V. Ortiz, A. F. Izmaylov, J. L. Sonnenberg, D. Williams-Young, F. Ding, F. Lipparini, F. Egidi, J. Goings, B. Peng, A. Petrone, T. Henderson, D. Ranasinghe, V. G. Zakrzewski, J. Gao, N. Rega, G. Zheng, W. Liang, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, K. Throssell, J. A. Montgomery, Jr, J. E. Peralta, F. Ogliaro, M. J. Bearpark, J. J. Heyd, E. N. Brothers, K. N. Kudin, V. N. Staroverov, T. A. Keith, R. Kobayashi, J. Normand, K. Raghavachari, A. P. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, J. M. Millam, M. Klene, C. Adamo, R. Cammi, J. W. Ochterski, R. L. Martin, K. Morokuma, O. Farkas, J. B. Foresman and D. J. Fox, *Gaussian 16, Revision C.01*, Gaussian Inc, Wallingford CT, 2019.
- 27 W. Kohn, Nobel Lecture: Electronic structure of matter-wave functions and density functionals, *Rev. Mod. Phys.*, 1999, **71**(5), 1253–1266.
- 28 D. Baric and Z. B. Maksic, Atomic Additivity of the Correlation Energy in Molecules by the DFT-B3LYP Scheme, *J. Phys. Chem. A*, 2003, **107**(51), 11577–11586.
- 29 D. Andrae, U. Haussermann, M. Dolg and H. Stoll, Energy-adjusted *ab initio* pseudopotentials for the second and third row transition elements, *Theor. Chim. Acta*, 1990, **77**(2), 123–141.
- 30 T. Lu, A Comprehensive Electron Wavefunction Analysis Toolbox for Chemists, Multiwfn, *J. Chem. Phys.*, 2024, **161**(8), 42.
- 31 J. S. Murray and P. Peter, The electrostatic potential: an overview, *Wiley Interdiscip. Rev. Comput. Mol. Sci.*, 2011, **1**(2), 153–163.
- 32 I. Mayer, Bond order and valence indices: A personal account, *J. Comput. Chem.*, 2007, **28**(1), 204–221.
- 33 S. J. Zheng, X. Y. Li and L. X. Mo, Introduction to Electron Localization Function(ELF)—A New Chemical Bond Model, *Mosc. Univ. Chem. Bull.*, 2010, **73**(3), 235–240.
- 34 T. LU and F. W. Chen, The meaning and functional form of the electron localization function, *Acta Phys.-Chim. Sin.*, 2011, **27**(12), 2786–2792.
- 35 J. Zhang, H. B. Yang, D. Zhou and B. Liu, Adsorption Energy in Oxygen Electrocatalysis, *Chem. Rev.*, 2022, **122**(23), 17028–17072.



- 36 S. Kozuch and S. Shaik, How to conceptualize catalytic cycles? The Energetic Span Model, *Acc. Chem. Res.*, 2011, **44**(12), 101–110.
- 37 S. Kozuch and S. Shaik, Kinetic-Quantum Chemical Model for Catalytic Cycles: The Haber-Bosch Process and the Effect of Reagent Concentration, *J. Phys. Chem. A*, 2008, **112**(26), 6032–6041.
- 38 K. U. Ansari, D. Borah, A. Rasamsetty, P. Kumar, M. Shanmugam, G. Rajaraman and M. Shanmugam, Probing the Slow Relaxation of Magnetization of a Square Planar Cobalt Complex with Doublet Ground State, *Chem. –Asian J.*, 2025, **20**(10), e202401798.
- 39 R. A. Miranda-Quintana, N. Adebar, M. Schulze and J. Smiatek, Sabatier Principle Revisited: The Role of Electronic Properties in Simple Catalytic Reactions, *J. Phys. Chem. C*, 2025, **129**(21), 9926–9934.
- 40 C. Molinet-Chinaglia, S. Shafiq and P. Serp, Low Temperature Sabatier CO₂ Methanation, *ChemCatChem*, 2024, **16**(24), e202401213.
- 41 S. Hu and W. X. Li, Sabatier principle of metal-support interaction for design of ultrastable metal nanocatalysts, *Science*, 2021, **374**, 1360–13655.
- 42 O. Elmutasim, L. M. Maghrabi, D. S. Dhawale and K. Polychronopoulou, Engaging the Concepts of Bimetallicity and Mechanical Strain for N₂ Activation: A Computational Exploration, *ACS Appl. Mater. Interfaces*, 2024, **16**, 56254–56270.
- 43 D. Y. Hwang and A. M. Mebel, Reaction Mechanism of N₂/H₂ Conversion to NH₃: A Theoretical Study, *J. Phys. Chem. A*, 2003, **107**, 2865–2874.
- 44 R. Cheng, C. N. Cui and Z. X. Luo, Reduction of dinitrogen to ammonia on doped three-atom clusters Nb₂M(M=Sc to Cu & Y to Ag), *Rare Met.*, 2024, **43**(8), 3810–3818.

