


 Cite this: *RSC Adv.*, 2025, 15, 40730

# Synthesis and solvatochromism of a NIR II emissive amphiphilic aza-BODIPY dye and its application as a colorimetric and fluorometric probe in sequential detection of Cu<sup>2+</sup> and PO<sub>4</sub><sup>3-</sup>

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The amphiphilic aza-BODIPY **1** bearing two *N,N*-dialkyl-amino groups at the *para*-positions of the 1,7-phenyl and two amino groups modified by a pair of hydrophilic chains at the *para*-positions at the 3,5-phenyl of the boron-azadipyrromethene core was synthesized and characterized. UV/Vis/NIR spectroscopic measurements indicated that the S<sub>0</sub>–S<sub>1</sub> absorption band **1** reached up to 900 nm while its emission band reached up to 1200 nm. The solvatochromic properties of **1** was investigated by absorption and fluorescence spectroscopy and further analysed by the Lippert–Mataga equation. The results suggested a strong intramolecular charge transfer (ICT) effect of dye **1** stemming from the electron-donating amino groups, which was further confirmed by DFT calculation studies. Such ICT could be disrupted upon addition of Cu<sup>2+</sup> owing to the metal–ligand interaction between Cu<sup>2+</sup> and the amino moieties, leading to a hypsochromic shift of the absorption band concomitant with a significant decrease in the fluorescence intensity of **1**. Moreover, the spectroscopic properties of **1** were recovered upon subsequent addition of PO<sub>4</sub><sup>3-</sup> in [**1** + Cu<sup>2+</sup>] solution. Based on these observations, the dye **1** was applied successfully as a NIR colorimetric and fluorometric “turn-off” probe for selective detection of Cu<sup>2+</sup> (LOD = 2.1 μM). Subsequent fluorometric “turn-on” detection of PO<sub>4</sub><sup>3-</sup> by [**1** + Cu<sup>2+</sup>] (LOD = 3.0 μM) was accomplished in further studies.

 Received 8th September 2025  
 Accepted 18th October 2025

DOI: 10.1039/d5ra06752a

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## Introduction

In recent years, functional dyes with near-infrared absorption and emission (NIR dyes)<sup>1</sup> have received widespread attention owing to their attractive potential in a variety of applications such as agents for phototherapy,<sup>2</sup> organic photovoltaics,<sup>3</sup> as well as fluorescence probes.<sup>4</sup> For the latter, the common probes operated within the ultraviolet/visible (UV/Vis) range could be limited by the background fluorescence interference, light scattering, or photodamage of the samples.<sup>5</sup> In contrast, the fluorescent probes functioned in the NIR-I (650–950 nm) and NIR-II (1000–1700 nm) windows<sup>6</sup> exhibited unique advantages including reduced susceptibility to background interference, high signal-to-noise ratio, less scattering and deeper penetration depth, and negligible impairment for biological samples.<sup>7</sup> Accordingly, NIR fluorescent sensing probes based on cyanine,<sup>8</sup> perylene diimide,<sup>9</sup> rhodamine,<sup>10</sup> BODIPYs,<sup>11</sup> *etc.* have been developed and applied for the detection of various metal ions as well as small molecules.<sup>12</sup>

As an important class of NIR chromophore, BF<sub>2</sub>-azadipyrromethenes (aza-BODIPYs)<sup>13</sup> exhibited intensive absorption and emission bands, high fluorescence quantum yields, as well as excellent stability. Moreover, the π-conjugate structure of the aza-BODIPY is facile to be chemically modified by introducing heavy atoms,<sup>14</sup> incorporating strong electron-donating groups to form a push–pull electronic structure,<sup>15</sup> and ring-fusing to form expanded conjugate structures of the aza-BODIPY dyes.<sup>16</sup> Based on these methods, the optical properties of aza-BODIPYs can be adjusted and their absorption and emission bands can be effectively shifted toward longer wavelengths in comparison with the parent tetraphenyl-substituted aza-BODIPY (λ<sub>abs</sub> = 650 nm, λ<sub>em</sub> = 673 nm).<sup>17</sup> For examples, the aza-BODIPY dyes bearing electron-donating triphenylamine<sup>18</sup> and julolidine<sup>19</sup> units were synthesized, which exhibited emission maxima at 892 and 1060 nm, respectively. Hao, Jiao, and coworkers<sup>20</sup> developed strategies for the synthesis of thieno-, and naphtho-fused aza-BODIPY dyes with their emission peaks at 816 nm and 705–972 nm, respectively.

While tremendous progresses were made in the synthesis of NIR absorbing aza-BODIPYs, fluorescent probes based on this class of dyes were developed for detecting various analytes,<sup>21</sup> including metal ions (Cu<sup>2+</sup>, Hg<sup>+</sup>),<sup>22</sup> anions (NH<sub>4</sub><sup>+</sup>, F<sup>-</sup>, NO<sub>2</sub><sup>-</sup>),<sup>23</sup>

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and small molecules ( $\text{H}_2\text{O}_2$ ,  $\text{H}_2\text{S}$ ,  $\text{NO}$ ).<sup>24</sup> Noteworthy that in recent years some sequential ion-responsive systems that were capable of detecting two different analytes have been reported. For examples, Kim *et al.*<sup>25</sup> developed a colorimetric probe for sequential detection of  $\text{Cu}^{2+}$  and  $\text{CN}^-$ . In another work, Ren *et al.*<sup>26</sup> synthesized a self-assembling tetraphenylethylene derivative for sequential detection of  $\text{Fe}^{3+}$  and  $\text{F}^-$ . Nevertheless, the sequential ion-responsive properties and detection of different analytes have not been reported yet for aza-BODIPY dyes, to the best of our knowledge.

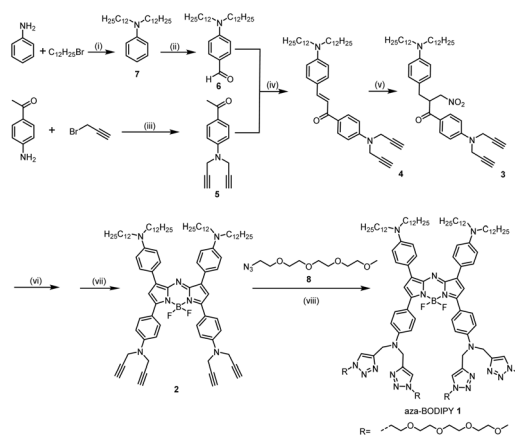
In the current work, we reported the synthesis, photo-physical properties, and sequential ion-responsive characteristics of an amphiphilic aza-BODIPY dye **1**. By introducing four electron-donating *N,N*-disubstituted amino groups at the *para*-positions of 1,3,5,7-phenyl substituents at the aza-BODIPY core, in polar solvent the dye **1** exhibited an emission band that reached the NIR-II window. In addition, the dye **1** bear two hydrophobic *N,N*-dodecyl groups at the 1,7-phenyl positions and four hydrophilic pendant containing oligo-ethylene glycol (OEG) chains. Such an amphiphilic feature of dye **1** offered good solubilities in both polar and nonpolar solvents, allowing the investigation of the solvatochromic properties of this dye in a wide range of solvents. Furthermore, the potential application of dye **1** as a colorimetric and fluorescent probe for the continuous detection of  $\text{Cu}^{2+}$  and  $\text{PO}_4^{3-}$  was also explored, which are known to be common environmental pollutants leading to health risks<sup>22a-c</sup> and deterioration of water quality.<sup>27</sup>

## Results and discussion

### Synthesis and spectroscopic properties of dye **1**

The synthetic route of the amphiphilic aza-BODIPY dye **1** was illustrated in Scheme 1. The chalcone derivative **4** was formed by the aldol condensation between the compounds **5** and **6**. The subsequent Michael addition between **4** and nitromethane was performed by using a 5-fold excess of nitromethane in ethanol at 80 °C for 12 h, affording intermediate **3**. Further condensation of **3** and ammonium acetate (15 equiv.) in refluxing *n*-butanol (110 °C, 12 h) gave the intermediate azadipyromethene, which was subsequently converted into aza-BODIPY **2** by chelation of  $\text{BF}_3 \cdot \text{Et}_2\text{O}$  in dried  $\text{CH}_2\text{Cl}_2$  under  $\text{N}_2$  atmosphere. In the last step, the amphiphilic aza-BODIPY dye **1** was obtained by attaching four oligo-ethylene glycol (OEG) hydrophilic pendants to the aza-BODIPY **2** with CuI-catalysed click reaction. The target dye **1** was characterized by  $^1\text{H}$  NMR,  $^{13}\text{C}$  NMR and MALDI-TOF mass spectroscopy (for details, see SI).

The UV/Vis/NIR absorption spectrum (black) of the amphiphilic dye **1** in  $\text{CHCl}_3$  ( $[\mathbf{1}] = 1.0 \times 10^{-6} \text{ M}$ ) was shown in Fig. 1a. An intensive band in NIR with a absorption maximum at 806 nm ( $\epsilon = 9.5 \times 10^4 \text{ M}^{-1} \text{ cm}^{-1}$ ) was observed, which could be assigned to the  $\text{S}_0\text{-S}_1$  transition of aza-BODIPY **1**.<sup>28</sup> This absorption maximum of **1** manifested a substantial bathochromic shift of 156 nm with respect to that of the parent tetraphenyl-substituted aza-BODIPY ( $\lambda_{\text{abs}} = 650 \text{ nm}$ ,  $\lambda_{\text{em}} = 673 \text{ nm}$ ),<sup>17</sup> which could be attributed to the introduction of the four *para*-dimethylaniline electron-donating into the molecular structure of dye **1**. In addition, the lower absorption band of dye **1**



Scheme 1 Synthesis of aza-BODIPY **1**. Reagents and conditions: (i)  $\text{K}_2\text{CO}_3$ , DMF, r.t., 80 °C, 24 h; (ii)  $\text{POCl}_3$ , DMF, r.t.; 79%; (iii)  $\text{K}_2\text{CO}_3$ , DMF, r.t., 24 h; 43%; (iv) *t*-BuOK, ethanol, r.t., 12 h; 41%; (v) *t*-BuOK,  $\text{CH}_3\text{NO}_2$ , ethanol, r.t., 80 °C, 24 h; 56%; (vi)  $\text{CH}_3\text{COONH}_4$ , *n*-butanol, 110 °C, 12 h; (vii)  $\text{BF}_3 \cdot \text{Et}_2\text{O}$ , dry  $\text{CH}_2\text{Cl}_2$ , DIEA, r.t., 12 h; 21%; (viii) CuI, DIEA,  $\text{CH}_2\text{Cl}_2$ ,  $\text{CH}_3\text{CN}$ , 60 °C, 8 h; 82%.

between 600 nm and 700 nm could be assigned to the  $\text{S}_0\text{-S}_2$  transition of this dye.

In Fig. 1a, the fluorescence spectrum (red) of aza-BODIPY **1** demonstrated a mirror relationship with its absorption spectrum, featuring a narrow-emission band with  $\lambda_{\text{em}} = 863 \text{ nm}$  and a Stokes shift of 57 nm. Meanwhile, a fluorescence quantum yield  $\Phi_{\text{fl}} = 0.16$  was determined for **1**. Moreover, the fluorescence lifetime of **1** was measured by time-resolved fluorescence spectroscopy (Fig. 1b). An average fluorescence lifetime  $\langle \tau \rangle = 4.00 \text{ ns}$  was obtained for dye **1** by fitting the decay curve with a biexponential model ( $\tau_1 = 1.39 \text{ ns}$ , 92.66% and  $\tau_2 = 37.00 \text{ ns}$ , 7.34%).

### Solvatochromic properties of dye **1**

To shed more light on the large bathochromic shift of the  $\text{S}_0\text{-S}_1$  absorption band of **1** with respect to that of the tetraphenyl-substituted the aza-BODIPY dye, the solvatochromic properties of dye **1** were investigated by absorption and emission spectroscopy in organic solvents of various polarity (Fig. 2 and Table 1).

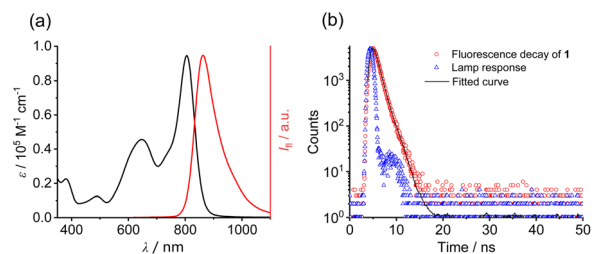
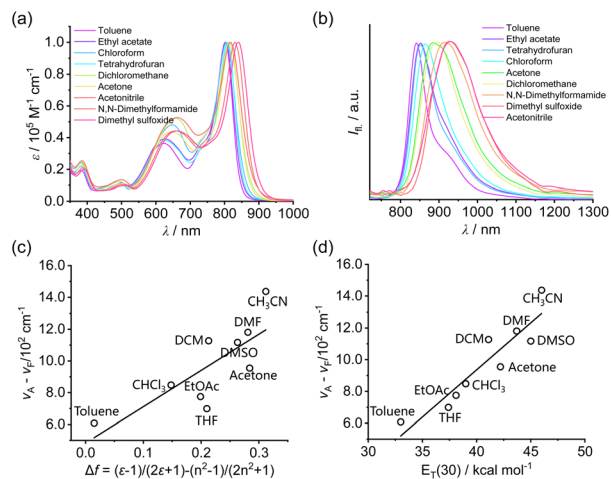


Fig. 1 (a) UV/Vis/NIR absorption and fluorescence spectra of dye **1** in  $\text{CHCl}_3$  ( $[\mathbf{1}] = 1.0 \times 10^{-6} \text{ M}$ ,  $\lambda_{\text{ex}} = 600 \text{ nm}$ ); (b) fluorescence decay spectra of dye **1** in  $\text{CHCl}_3$  ( $[\mathbf{1}] = 1.0 \times 10^{-6} \text{ M}$ ,  $\lambda_{\text{ex}} = 600 \text{ nm}$ ,  $\lambda_{\text{em}} = 820 \text{ nm}$ ).





**Fig. 2** UV/Vis/NIR absorption spectra (a) and fluorescence spectra (b) of dye **1** in different organic solvents; (c) Lippert equation of dye **1** in different organic solvents; (d) the relationship between the Stokes shift ( $v_A - v_F$ ) of dye **1** and the empirical polarity parameter  $E_T(30)$  of different organic solvents. ( $[1] = 1.0 \times 10^{-6}$  M,  $\lambda_{\text{ex}} = 700$  nm).

**Table 1** Spectral properties of dye **1** in different solvents, including chloroform ( $\text{CHCl}_3$ ), *N,N*-dimethylformamide (DMF), acetonitrile ( $\text{CH}_3\text{CN}$ ), tetrahydrofuran (THF), dimethyl sulfoxide (DMSO), dichloromethane ( $\text{CH}_2\text{Cl}_2$ ), ethyl acetate (EtOAc), toluene, acetone

Solvent	$\epsilon^a$	$n^b$	$\Delta f^c$	$\lambda_A^d/\text{nm}$	$\lambda_F^e/\text{nm}$	$(v_A - v_F)^f/\text{cm}^{-1}$
Toluene	2.4	1.496	0.015	801	842	608
EtOAc	6.0	1.372	0.199	801	854	775
$\text{CHCl}_3$	4.8	1.447	0.148	805	864	848
THF	7.6	1.407	0.210	805	853	699
$\text{CH}_2\text{Cl}_2$	8.5	1.424	0.213	813	895	1127
Acetone	20.7	1.359	0.284	816	885	955
$\text{CH}_3\text{CN}$	37.5	1.344	0.312	818	927	1437
DMF	36.7	1.431	0.281	829	919	1181
DMSO	46.7	1.478	0.263	840	927	1117

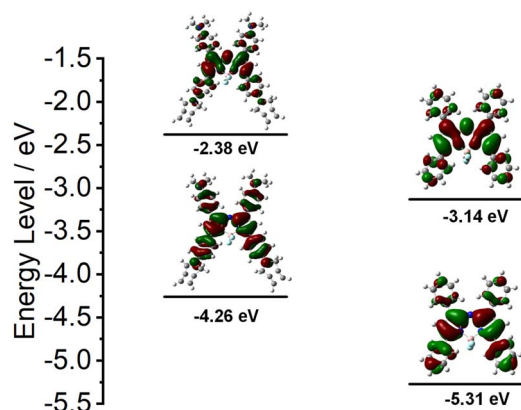
<sup>a</sup> Permittivity of solvent. <sup>b</sup> Refractive index of solvent. <sup>c</sup>  $\Delta f = [(\epsilon - 1)/(2\epsilon + 1)] - [(n^2 - 1)/(2n^2 + 1)]$ . <sup>d</sup> Absorption maxima. <sup>e</sup> Emission maxima. <sup>f</sup> Stokes shift.

As shown in Fig. 2a, the absorption maxima of dye **1** in different solvents were located between 801–840 nm, depending on the solvent polarity. In contrast to the absorption spectra, much larger shift of the emission maxima was observed in the fluorescence spectra of dye **1** (Fig. 2b) in different solvents. With increase in solvent polarity, the emission maxima exhibited a considerable red-shift from 842 nm (toluene) to 927 nm (DMSO), indicating a positive solvatochromic effect<sup>29</sup> of dye **1**.

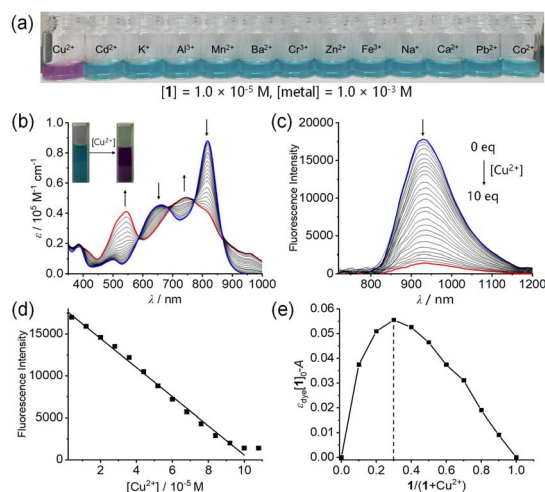
Based on the spectroscopic data summarized in Table 1, the solvatochromic properties of dye **1** were further analysed by Lippert–Mataga eqn (1),<sup>30</sup>

$$hc(v_A - v_F) = 2\Delta f \frac{(\mu_E - \mu_G)^2}{a^3} + C \quad (1)$$

where  $v_A$  and  $v_F$  were the absorption and emission maxima in wave numbers dye **1** ( $\text{cm}^{-1}$ ),  $\mu_G$  and  $\mu_E$  were the dipole moments



**Fig. 3** Distributions of HOMO and LUMO and energy levels of a model compound for **1** (left, the hydrophobic and hydrophilic chains in **1** were replaced by methyl groups) and tetraphenyl-substituted aza-BODIPY. DFT calculations were carried out on the Gaussian 09W software platform, using B3LYP as the method and 6-311G(d) as the base group.

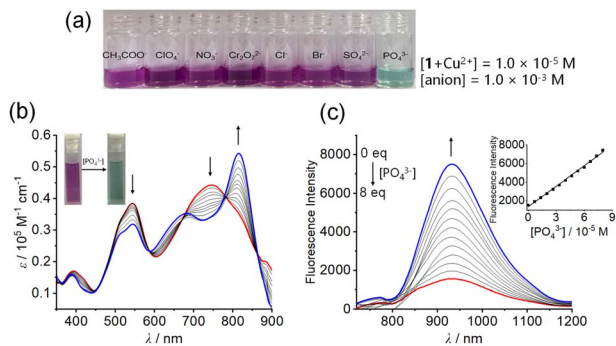


**Fig. 4** (a) The colours of the solutions after adding different metal ions. ( $[1] = 1.0 \times 10^{-5}$  M,  $[\text{metal}] = 1.0 \times 10^{-3}$  M); (b) UV/Vis absorption spectra (Inset: Colour change of the solution after adding  $\text{Cu}^{2+}$ ); (c) fluorescence spectra of dye **1** in  $\text{CH}_3\text{CN}/\text{H}_2\text{O}$  ( $v/v, 10 : 1$ ); (d) plot of the fluorescence intensity of the system with different concentrations of  $\text{Cu}^{2+}$ ; (e) Job's plot of dye **1** and  $\text{Cu}^{2+}$  in  $\text{CH}_3\text{CN}$ . (The absorbances ( $A$ ) of the solutions at 818 nm were used for analysis.)

of the ground state and excited state respectively,  $h$  was the Planck's constant,  $c$  was the velocity of light,  $a$  was the radius of the cavity where the fluorophore was located (Onsager radius), and  $\Delta f$  was the solvent parameter, which was defined as  $[(\epsilon - 1)/(2\epsilon + 1)] - [(n^2 - 1)/(2n^2 + 1)]$ .

As shown in Fig. 2c, we observed an approximate linear relationship between the Stokes shifts ( $v_A - v_F$ ) of dye **1** and the solvent parameters ( $\Delta f$ ). Based on the slope  $(2/hc)[(\mu_E - \mu_G)^2/a^3]$ <sup>31</sup> obtained by linear fitting and a cavity radius  $a = 0.51$  nm of the chromophore, which was estimated according to the molecular modelling, the change of dipole moment between the excited state and the ground state ( $\Delta\mu_{\text{eg}}$ ) of dye **1** was calculated to be 1.74 D, indicating that the dipole moment of





**Fig. 5** (a) The colours of the solutions after adding different anions. ( $[1 + \text{Cu}^{2+}] = 1.0 \times 10^{-5} \text{ M}$ ,  $[\text{anion}] = 1.0 \times 10^{-3} \text{ M}$ ); (b) UV/Vis absorption spectra (Inset: Colour change of the solution after adding  $\text{PO}_4^{3-}$ ); (c) fluorescence spectra of this system in  $\text{CH}_3\text{CN}/\text{H}_2\text{O}$  (v/v, 50 : 9) (Inset: Plot of the fluorescence intensity of the system with different concentrations of  $\text{PO}_4^{3-}$ ).

**Table 2** Comparisons between dye 1 and literature-reported probes A–G (mostly BODIPY/aza-BODIPY-based, see Fig. S33 for chemical structures) for detection of  $\text{Cu}^{2+}$  or  $\text{PO}_4^{3-}$  or for sequential ion-detection

Probe	$\lambda_{\text{abs}}^a$	$\lambda_{\text{em}}^b$	$\Phi_{\text{fl}}^c$	$\text{LOD}_{\text{Cu}^{2+}}$	$\text{LOD}_{\text{anion}}$
<b>1</b>	818	927	0.16	2.1 $\mu\text{M}$	3.0 $\mu\text{M}$
<b>A<sup>d</sup></b>	450	N.A.	N.A.	0.9 $\mu\text{M}$	( $\text{CN}^-$ )
<b>B<sup>e</sup></b>	N.A.	420	N.A.	N.A.	19.69 mM
<b>C<sup>f</sup></b>	503	580	0.34	6 $\mu\text{M}$	N.A.
<b>D<sup>g</sup></b>	686	720	N.A.	N.A.	N.A.
<b>E<sup>h</sup></b>	680	717	N.A.	0.2 $\mu\text{M}$	N.A.
<b>F<sup>i</sup></b>	690	725	N.A.	3.75 ppb	N.A.
<b>G<sup>j</sup></b>	784	840	0.29	1.38 $\mu\text{M}$	N.A.

<sup>a</sup> Absorption Maximum. <sup>b</sup> Emission maximum. <sup>c</sup> Fluorescence quantum yield. <sup>d</sup> Ref. 25. <sup>e</sup> Ref. 27. <sup>f</sup> Ref. 41. <sup>g</sup> Ref. 11. <sup>h</sup> Ref. 22a. <sup>i</sup> Ref. 22b. <sup>j</sup> Ref. 15b.

the excited state of the dye was greater than that of the ground state. This result ( $\mu_{\text{E}} > \mu_{\text{G}}$ ) could be explained by the formation of intramolecular charge transfer (ICT) state upon excitation of dye 1.<sup>32</sup> Such an ICT state could be stabilized upon increase in the solvent polarity. Accordingly, a positive solvatochromic effect was displayed by dye 1.

In addition, while the empirical solvent polarity parameter  $E_{\text{T}}(30)$ <sup>33</sup> was used to study the relationship between the Stokes shifts of dye 1 and solution polarity (Fig. 2d), a correlation coefficient  $r = 0.93$  was obtained, indicating that the correlation was significantly improved as compared with that using  $\Delta f$ .

In order to further elucidate the ICT characteristics of dye 1, density functional theory (DFT) calculations were performed for a model compound of dye 1 (in which the side chains were replaced by methyl groups) as well as the tetraphenyl-substituted aza-BODIPY (Fig. 3). The results suggested that the HOMOs of 1 were delocalized over the entire molecular framework. In contrast, the LUMOs of 1 were mainly located on the aza-BODIPY core with an obvious contribution of the *meso*-N atom. Such separation between HOMO and LUMO implied that

ICT might take place between the electron-donating amino groups and the electron-accepting aza-BODIPY core upon excitation of dye 1. In contrast, tetraphenyl-substituted aza-BODIPY displayed a less extensive HOMO than the model compound. In addition, the calculation pointed to a smaller band gap of dye 1 than that of tetraphenyl-substituted aza-BODIPY, which was in agreement with the results of spectroscopic studies.

### Colorimetric and fluorometric “turn-off” probe for $\text{Cu}^{2+}$

Since it has been reported that the bis(1,2,3-triazole) amino group was effective ligand for metal ions,<sup>34</sup> the application potential of dye 1 as sensor for metal ion detection was investigated. Accordingly, solutions of various metal ions ( $1.0 \times 10^{-3} \text{ M}$ ) in water (10 equiv.) were added to dye 1 solutions ( $1.0 \times 10^{-5} \text{ M}$ ) in  $\text{CH}_3\text{CN}$ . The latter was not only a good solvent for dye 1 but also miscible with water and used widely in ion-detection studies.<sup>11,15b,35</sup> While the colour of the dye solution changed remarkably from blue to purple upon adding of  $\text{Cu}^{2+}$ , no colour changes were observed by naked eyes upon adding of the other metal ions, including  $\text{Cd}^{2+}$ ,  $\text{K}^+$ ,  $\text{Al}^{3+}$ ,  $\text{Mn}^{2+}$ ,  $\text{Ba}^{2+}$ ,  $\text{Cr}^{3+}$ ,  $\text{Zn}^{2+}$ ,  $\text{Fe}^{3+}$ ,  $\text{Na}^+$ ,  $\text{Ca}^{2+}$ ,  $\text{Pb}^{2+}$ ,  $\text{Co}^{2+}$  (Fig. 4a), implying that dye 1 could be used for selective detection of  $\text{Cu}^{2+}$ . This was further confirmed by the absorption and fluorescence spectroscopic measurements (Fig. S23). In the presence of  $\text{Cu}^{2+}$ , significant spectral changes were observed in the absorption and fluorescence spectra of dye 1. In contrast, negligible spectral responses of the dye 1 solutions were detected for other ions. Thus, the aza-BODIPY 1 could serve as a highly selective colorimetric and NIR fluorometric “turn-off” probe<sup>36</sup> for  $\text{Cu}^{2+}$ .

To quantitatively elucidate the binding between dye 1 and  $\text{Cu}^{2+}$ , UV/Vis and fluorescence spectroscopic titration studies were performed. Upon gradual addition of up to 10 equiv. of  $\text{Cu}^{2+}$ , gradual lowering of the  $\text{S}_0\text{-S}_1$  (around 818 nm) and  $\text{S}_0\text{-S}_2$  band (around 659 nm) was observed (Fig. 4b). The absorption maximum was hypsochromically shifted from 818 nm to 745 nm, indicating that the ICT effect in dye 1 could be disrupted by the metal–ligand interactions between  $\text{Cu}^{2+}$  and the bis (1,2,3-triazole) amino groups. Meanwhile, new absorption bands gradually raised at around 745 nm and 544 nm. Such absorption spectra changes were in agreement with the colour change of the dye solution upon adding the  $\text{Cu}^{2+}$ . Moreover, decrease in the fluorescence intensity at around 955 nm was observed upon the addition of  $\text{Cu}^{2+}$  (Fig. 4c). The plot of fluorescence intensity *versus* the concentration of  $\text{Cu}^{2+}$  exhibited a good linear relationship ( $R^2 = 0.9954$ ), demonstrating that dye 1 was capable of detecting  $\text{Cu}^{2+}$  quantitatively. The Stern–Volmer plot<sup>37</sup> based on the fluorescence spectra (Fig. S24) of dye 1 gave a quenching constant  $K_{\text{SV}} = 3.7 \times 10^4 \text{ M}^{-1}$  by linear data fitting. Based on the change of fluorescence spectra at 929 nm, the limit of detection (LOD)<sup>38</sup> for  $\text{Cu}^{2+}$  was determined to be 2.1  $\mu\text{M}$  (Fig. S25).

To verify the binding stoichiometry between the dye 1 and  $\text{Cu}^{2+}$ , Job's plot analysis was performed for the absorbance variations at 818 nm (Fig. 4e) and the result unambiguously pointed to a 1 : 2 binding stoichiometry of 1 and  $\text{Cu}^{2+}$ . Thus, each of the bis(1,2,3-triazole) amino moieties in dye 1 was



bound to one  $\text{Cu}^{2+}$  and a bimetallic complex of **1** was formed. This was further confirmed by the mass spectroscopic (MALDI-TOF) measurement of the complex (Fig. S26), which indicated the formation of species  $\mathbf{1} + 2\text{Cu}^{2+}$  ( $m/z$  610.5, calculated: 610.5) and  $\mathbf{1} + \text{Cu}^{2+}$  ( $m/z$  1189.9, calculated: 1189.3). The binding constant of dye **1** and  $\text{Cu}^{2+}$  could be estimated by using eqn (2),<sup>39</sup>

$$\frac{[\text{H}]_0[\text{G}]_0^2}{\Delta A} = \frac{1}{\Delta \epsilon_a K} + \frac{[\text{G}]_0([\text{G}]_0 + 4[\text{H}]_0)}{\Delta \epsilon_a} \quad (2)$$

where  $[\text{H}]_0$  and  $[\text{G}]_0$  were the initial concentrations of **1** (host) and  $\text{Cu}^{2+}$  (guest), respectively, and  $\Delta A$  and  $\Delta \epsilon_a$  represented the changes of absorbance and molar absorption coefficient after complexation, respectively. By linear least-squares fitting with eqn (2), an overall binding constant  $K = 2.0 \times 10^9 \text{ M}^{-2}$  was obtained (Fig. S27).

In addition, competitive binding test was conducted to evaluate the  $\text{Cu}^{2+}$  selectivity of **1** by introducing 10 equiv. excess of competing cations into the  $\mathbf{1} + \text{Cu}^{2+}$  complex (1:2 stoichiometry), including  $\text{Cd}^{2+}$ ,  $\text{K}^+$ ,  $\text{Al}^{3+}$ , etc. The fluorescence spectra of **1** (Fig. S28) indicated that, the fluorescence intensity of the solutions remained unchanged before and after the addition of other metal ions, indicating that  $\text{Cu}^{2+}$  in the complex could not be replaced by other metal ions and had a good anti-interference capability for the detection of  $\text{Cu}^{2+}$ .

### Fluorometric “turn-off-on” probe for sequential detection of $\text{PO}_4^{3-}$

Since the dye **1** was capable of chelating with  $\text{Cu}^{2+}$ , it was attractive to further investigate the response properties and potential of the  $[\mathbf{1} + \text{Cu}^{2+}]$  system in sequential detection of anions.<sup>25</sup> Accordingly, a variety of common anions in water including  $\text{PO}_4^{3-}$ ,  $\text{ClO}_4^-$ ,  $\text{Cl}^-$ ,  $\text{Br}^-$ ,  $\text{NO}_3^-$ ,  $\text{Cr}_2\text{O}_7^{2-}$ ,  $\text{SO}_4^{2-}$ ,  $\text{CH}_3\text{COO}^-$  were introduced into solution of  $[\mathbf{1} + \text{Cu}^{2+}]$ . Upon addition of  $\text{PO}_4^{3-}$ , the colour of the dye **1** solution was changed to be blue-green. Meanwhile, no colour changes were observed for the solution upon addition of the other anions (Fig. 5a), suggesting a selective response of  $[\mathbf{1} + \text{Cu}^{2+}]$  to  $\text{PO}_4^{3-}$ , which was corroborated by monitoring the absorption and fluorescence spectral changes in the presence of various of anions (Fig. S29).

As shown in Fig. 5b and c, the absorption bands of dye **1** were increased with the increasing concentration of  $\text{PO}_4^{3-}$  while the fluorescence intensity of **1** was increased drastically. This result indicated that  $\text{PO}_4^{3-}$  might have a stronger binding capability to  $\text{Cu}^{2+}$  than that of **1**, leading to the disassembly of  $[\mathbf{1} + \text{Cu}^{2+}]$  species and the recovery of fluorescence of **1**. In addition, the fluorescence intensity of the system exhibited a good linear relationship with the concentration of  $\text{PO}_4^{3-}$  ( $R^2 = 0.9984$ , Fig. 5c inset). Moreover, the LOD for the detection of  $\text{PO}_4^{3-}$  was determined to be  $3.0 \mu\text{M}$  (Fig. S30). To further verify the specificity of  $\text{PO}_4^{3-}$  detection, competition experiments were conducted (Fig. S31) and the results indicated that there was no significant change in the fluorescence intensity of the system in the presence of other anions, demonstrating that this “turn-off-on”<sup>40</sup> probing system for  $\text{PO}_4^{3-}$  detection was effective in complex ionic environments.

The reversibility of the probing system after adding  $\text{Cu}^{2+}$  and  $\text{PO}_4^{3-}$  was evaluated. Upon further adding  $\text{Cu}^{2+}$  and  $\text{PO}_4^{3-}$ , the absorption bands of dye **1** and  $[\mathbf{1} + \text{Cu}^{2+}]$  was further recovered (Fig. S32), indicating the reversible ion-responsive behaviour of dye **1**. In addition, parameters of dye **1** and representative probes A–G for detection of  $\text{Cu}^{2+}$  or  $\text{PO}_4^{3-}$  or for sequential ion-detection (mostly BODIPY/aza-BODIPY-based, see Fig. S33 for chemical structures) were listed in Table 2. These data indicated that dye **1** exhibited moderate LODs of  $\text{Cu}^{2+}$  and  $\text{PO}_4^{3-}$  with the other probes. Meanwhile, aza-BODIPY **1** displayed the longest wavelength of absorption (818 nm) and emission maximum (927 nm), underscoring its better applicability as ion probes for biological systems.

## Conclusions

In summary, a new amphiphilic NIR aza-BODIPY dye **1** bearing four hydrophobic alkyl chains and four hydrophilic oligo-ethylene glycol chains was synthesized and characterized. By introducing four *N,N*-double substituted amino groups as electron donating groups into dye **1**, it exhibited attractive NIR spectroscopic properties with absorption wavelength up to 950 nm and emission wavelength up to 1200 nm. A positive solvatochromism was observed for dye **1**. Further analysis of the solvent-dependent spectroscopic properties by Lippert–Mataga equation and DFT calculation studies indicated that the large bathochromic spectral shift of **1** with respect to the parent tetraphenyl-substituted aza-BODIPY could be ascribed to the strong ICT effect stemming from the electron-donating amino groups. Further observations of solution colour changes by naked eye and spectroscopic measurements implied that the ICT process in dye **1** could be disrupted upon addition of  $\text{Cu}^{2+}$  and recovered upon subsequent addition of  $\text{PO}_4^{3-}$ . Accordingly, the dye **1** was applied as a NIR colorimetric and fluorometric “turn-off” probe for selective detection of  $\text{Cu}^{2+}$  with a LOD of  $2.1 \mu\text{M}$ . In addition, the complex  $[\mathbf{1} + \text{Cu}^{2+}]$  was demonstrated to be a fluorometric “turn-on” probe for the detection of  $\text{PO}_4^{3-}$  with a LOD of  $3.0 \mu\text{M}$ .

## Experimental section

All the reagents and solvents in the synthesis were obtained from commercial suppliers and were not purified. All solvents used for spectral measurements were chromatographically pure and required no further purification. Details of the synthesis and characterization of intermediates 2–8 were provided in SI.

### General methods

$^1\text{H}$  and  $^{13}\text{C}$  NMR spectra were recorded on Bruker AVANCE III HD 400 MHz or JEOL JNM ECZ600R spectrometer at room temperature. Chemical shifts ( $\delta$ ) were given in ppm relative to  $\text{CDCl}_3$  (7.26 ppm for  $^1\text{H}$  and 77 ppm for  $^{13}\text{C}$ ) to internal TMS. Mass spectra was measured by Autoflex III matrix-assisted laser desorption ionization time-of-flight (MALDI-TOF) mass spectrometry with  $\alpha$ -cyano-4-hydroxycinnamic acid as the matrix. UV-Vis-NIR absorption spectroscopic measurements were



performed on UV-3600 (Shimadzu) or Carry 300 (Agilent) spectrophotometer. All measurements were made at 25 °C, using 10 × 10 mm cuvettes. Molar absorption coefficient  $\epsilon$  was calculated according to Lambert–Beer's law.

### Fluorescence spectroscopic measurement

The fluorescence spectra, fluorescence quantum yield of **1**, and fluorescence lifetimes were measured by a FLS 1000 spectrofluorometer (Edinburgh). All the fluorescence spectra were corrected. The fluorescence quantum yield of dye **1** was measured by using a built-in integrating sphere in the spectrofluorometer. The fluorescence lifetime measurements were performed with a time-correlated single photon counting setup consisting of a pulsed light emitting diode laser excitation source ( $\lambda_{\text{ex}} = 635$  nm). The instrument response was collected by scattering the excited light of a dilute, aqueous suspension of colloidal silica (Ludox). The lifetime decay curves were analysed using the software supplied with the instrument. The quality of the data fitting was evaluated by analysis of  $\chi^2$  (0.9–1.1) as well as by inspection of residuals and the autocorrelation function.

### Synthesis of aza-BODIPY **1**

Compound **2** (0.138 g, 0.1 mmol) was dissolved in the mixed solvent of  $\text{CH}_2\text{Cl}_2/\text{CH}_3\text{CN}$  (40 mL, 1/1, v/v), and then  $\text{CuI}$  (0.019 g, 0.1 mmol), DIEA (18 drops), and compound **8** (0.198 g, 0.8 mmol) were added. The resultant reaction mixture was stirred for 8 h at 60 °C. After the reaction was completed (monitored by TLC), the solvent was rotary-evaporated. The oily residue was dissolved in  $\text{CH}_2\text{Cl}_2$  and washed with water (3 × 50 mL). The combined organic layers were dried over  $\text{MgSO}_4$  and concentrated. Further purification by column chromatography (silica gel, 300–400 mesh) using  $\text{MeOH}/\text{CH}_2\text{Cl}_2$  (1/15, v/v) as the eluent gave the aza-BODIPY dye **1** as a blue-black solid (0.175 g, 82%).  $^1\text{H}$  NMR (600 MHz, chloroform-*d*):  $\delta = 8.03$  (d,  $J = 8.4$  Hz, 8H), 7.67 (s, 4H), 6.93 (d,  $J = 7.7$  Hz, 4H), 6.81 (s, 2H), 6.67 (d,  $J = 8.6$  Hz, 4H), 4.80 (s, 8H), 4.50 (s, 8H), 3.83 (d,  $J = 4.9$  Hz, 8H), 3.54–3.30 (m, 68H), 1.63 (d,  $J = 7.3$  Hz, 8H), 1.33 (d,  $J = 27.9$  Hz, 72H), 0.87 (t,  $J = 7.0$  Hz, 12H).  $^{13}\text{C}$  NMR (101 MHz, chloroform-*d*):  $\delta = 55.35, 149.08, 148.52, 144.67, 141.57, 131.18, 130.74, 123.31, 121.21, 120.52, 113.98, 112.61, 111.32, 71.84, 70.57, 70.45, 70.37, 69.45, 58.95, 53.46, 51.15, 50.31, 46.60, 31.95, 31.44, 30.19, 29.77, 29.72, 29.68, 29.62, 29.39, 27.29, 22.71, 14.15$ . MS (MALDI-TOF):  $m/z$  calculated for  $\text{C}_{128}\text{H}_{206}\text{BF}_2\text{N}_{19}\text{O}_{16}$ ,  $[\text{M}]^+ = 2316.0$ , found: 2316.4. Elemental analysis: calculated for  $\text{C}_{128}\text{H}_{206}\text{BF}_2\text{N}_{19}\text{O}_{16}$ : C 66.38%, H 8.97%, N 11.49%; found: C 66.65%, H 8.50%, N 10.85%. UV-Vis-NIR absorption (in  $\text{CHCl}_3$ ):  $\lambda_{\text{max}}(\epsilon) = 805$  (95 000), 648 (46 000), 430 (13 000), 380 (23 000  $\text{M}^{-1} \text{cm}^{-1}$ ); fluorescence (in  $\text{CHCl}_3$ ):  $\lambda_{\text{max}} = 863$  nm; fluorescence quantum yield (in  $\text{CHCl}_3$ ):  $\Phi_{\text{fl}} = 0.16$ .

### Conflicts of interest

There are no conflicts to declare.

### Data availability

The data supporting this article have been included as part of the supplementary information (SI). Supplementary information:  $^1\text{H}$  NMR and  $^{13}\text{C}$  NMR spectra, absorption and fluorescence spectra, and other results. See DOI: <https://doi.org/10.1039/d5ra06752a>.

### Acknowledgements

The authors acknowledge the financial support of the National Natural Science Foundation of China (Grant No. 92056115).

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