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Ternary nanocomposite ReMIL-CN (ReSe₂@MIL-53(Fe)@g-C₃N₄) for energy storage and electroactive integrated H₂O₂ sensor application

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Integrated electrochemical systems capable of simultaneous charge storage and biochemical sensing response have gained prominence in modern materials science and biomedical engineering by offering multifunctionality for next-generation smart devices. The present comprehensive study is focused on the synthesis and characterization of a ternary nanocomposite material, focusing on its applications in integrated electrode systems for BHSCs and hydrogen peroxide (H₂O₂) sensing. The synthesized ReMIL-CN (ReSe₂@MIL-53(Fe)@g-C₃N₄) nanocomposite exhibits an average crystallite size of 27.27 nm, a specific surface area of 154 m² g⁻¹, and a specific pore volume of 0.022 cm³ g⁻¹. The ReMIL-CN electrode offers maximum specific capacities (Q_s) of 1925.5 C g⁻¹ in a three-electrode testing setup and 423 C g⁻¹ in a ReMIL-CN//AC full-cell setup. The values of the oxidation and reduction diffusion coefficients (D_{oxidation} and D_{reduction}) of the ReMIL-CN compound are 4.59 × 10⁻⁸ m² s⁻¹ and 1.45 × 10⁻⁸ m² s⁻¹, respectively. The highest obtained values of the energy density (E_d) and power density (P_d) are 50 Wh kg⁻¹ and 2060 W kg⁻¹, respectively. The ReMIL-CN//AC hybrid device demonstrates high stability response, with 95% capacity retention, 90% coulombic efficiency, and 90% and 84% charge–discharge time retention over a prolonged period of 10 000 repeated cycles. The measured *b*-values of 0.59, 0.68, and 0.70 suggest the hybrid-natured operational mechanism of the ReMIL-CN//AC device. Moreover, the present composite-based electrode exhibits a sensitivity of S = 0.185 μA μM⁻¹ cm² with a linear operational range of 50–1000 μM for the H₂O₂ detection system. The device maintains 95% signal stability after 1000 cycles, with an LOD of 0.1 μM, a linear regression value of R² = 0.998, and a minimal response time of less than 3 seconds, demonstrating the excellent selectivity of the device. The outstanding electrochemical charge storage and sensing capabilities of this ternary nanomaterial enable its integration into real-time, self-powered bioelectrochemical monitoring for next-generation portable and wearable bioelectronics.

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1 Introduction

Efficient electrochemical energy storage and conversion technologies are pivotal for addressing the rising global energy requirements and environmental issues.^{1–4} The foremost challenge in this field is the development of high-energy- and high-power-density delivery systems with long-term operational periods.^{5–7} Supercapacitors (SCs) and batteries are widely accepted as safe, clean, and green electrochemical entities for power sector applications.^{8–11} However, batteries inherently suffer from slower charge–discharge rates, while supercapacitors have comparatively lower energy storage capacities. The synergistic integration of these two devices into a single hybrid system helps overcome their limitations.^{12,13} Such hybrid systems can facilitate both rapid energy release and high energy retention along with efficient lifespans and cost-effectiveness, making these battery-hybrid supercapacitors (BHSCs) highly suitable for efficient energy storage systems.^{14–17} Integrated



detection systems based on these electrochemical BHSC devices have also emerged as powerful analytical tools for food and water quality monitoring and medical diagnostics.^{18–20} The capabilities of these BHSC devices for holding high energy density *via* a rapid charge–discharge process enable them to achieve efficient, real-time, and portable electrochemical sensing response. Such electrochemical systems can be used for contaminant detection in the food and water quality sectors to ensure public health safety and in medical field applications, where these hybrid systems enable sensitive, selective, and trace-level detection of different biomarkers and chemicals, along with efficient power storage and supply.^{21–24} Hydrogen peroxide (H₂O₂) is a commonly used chemical in household, medical, and industrial applications. However, it can also cause severe health risks depending on its concentration and exposure method. Therefore, early and trace-level detection and proper handling of H₂O₂ are essential to minimize these risks.^{25,26}

The pursuit of enhanced performance in hybrid energy devices has directed researchers towards the development of advanced electrode materials.^{27–29} A variety of composite materials combining metallic, organic, and inorganic components, such as MOFs, TMDs, TMOs, TMSs, HKUSTs, conductive polymers, and MILs, have drawn significant attention owing to their unique structural features, high surface areas, and tunable porosities, which make them ideal for dynamic charge storage applications.^{30,31} These materials enhance charge accommodation and transport, increase the specific energy and power density, and improve the charge–discharge capabilities, leading to high energy storage and rapid pulsation response applications.^{32–34} The advancements in electrode materials have also enhanced the sensitivity, selectivity, and stability of these electrode-based sensors.³⁵ Integrating detection systems with hybrid energy storage mechanisms improves operational longevity and supports self-powered sensing applications, leading to multifunctional electrochemical sensing systems for remote and point-of-care diagnostics.^{36,37}

Subhani, Karamat, *et al.* stated that the MnO₂-decorated carbon fiber has the highest specific capacitance of 1.71 F g^{−1} and an energy density of 3.8 Wh kg^{−1} with a 123.6 MPa mechanical strength in supercapacitor-illuminated LEDs.³⁸ Xia, Changlei, *et al.* reported a CoFe₂O₄/GNRs electrode offering specific capacitances of 922 F g^{−1} and 487.85 F g^{−1} in an asymmetric supercapacitor cell system, with the highest power and energy densities of 6730.76 W kg^{−1} and 32.8 Wh kg^{−1}, respectively.³⁹ Zhang, Fan, *et al.* presented a staged intercalation–deintercalation mechanism in the LiFePO₄-hybrid-graphite electrode, with an energy density of 176.7 Wh kg^{−1}, good capacity, and efficient stability for over 3500 consecutive repeated cycles.⁴⁰ Hu, Pu, *et al.* presented a novel sodium-based NaV₃(PO₄)₃/C electrode for battery testing, exhibiting a capacity of 146 mAh g^{−1} in a Na half-cell study and 80% capacity retention after 1000 cycles in a symmetric electrode full-cell configuration.⁴¹ Sakthivel, P., *et al.* reported the development of a MoS₂-Bi₂S₃@CNT-alkaline electrolytic system having a specific capacitance of 1338 F g^{−1}, a 99.5% stable columbic efficiency response over 10 000 cycles, and 60% and 34.6%

retention stability responses over 2000 and 10 000 cycles, respectively.⁴² Cui, Xiaosha, *et al.* reported aqueous MnO₂ electrode-based AEESD energy devices exhibiting a high energy density of 0.12 mWh cm^{−2} and a long lifespan, with a capacitance of 1551 mF cm^{−2}, and 89.63% stable capacitance retention.⁴³ Sahoo, Ramkrishna, *et al.* reported a highly stable rGO@VO₂ heterostructured anode-based BSH with AC@CC as the cathode, which exhibited the specific capacity, energy, and power of 1214 mAh g^{−1}, 126.7 Wh kg^{−1}, and 10 000 W kg^{−1}, respectively.⁴⁴ Alam, Faiz, *et al.* reported the Co₃(PO₄)₂@PANI hybrid system possessing a power density of 6027 W kg^{−1}, a 97.6% cyclic stability, an 87% columbic efficiency, and an energy density of 7.1 Wh kg^{−1} in a supercapattery-type assembly.⁴⁵ Nguyen, P. *et al.* reported TiO₂/TiS₂ for lithium–sulfur batteries with an energy density of 331 Wh kg^{−1} and good cyclic response.⁴⁶ Ahmad, Awais, *et al.* reported aqueous hybrid supercapacitors (AHSCs) using CCO (CoCr₂O₄) electrodes, exhibiting a specific capacitance of 2951 F g^{−1}, a minimum charge transfer resistance (*R*_{ct}) value of 0.135 Ω, and 98.7% stable capacitance retention response.⁴⁷ He, Jiarui, *et al.* reported Na–S batteries having an energy density of 1274 Wh kg^{−1} and a discharge capacity of 1081 mAh g^{−1}.⁴⁸ Zhang, Fan, *et al.* reported a graphene-enhanced Fe₃O₄ electrode with a 3D graphene electrode in a supercapacitor–battery hybrid system, which had the highest energy and power density values of 147 Wh kg^{−1} and 2587 W kg^{−1}, respectively.⁴⁹ Zhang, Yang, *et al.* reported a lithium-ion capacitor-based system comprising AC negative and graphite positive electrodes, which exhibited the highest energy and power densities of 200 Wh kg^{−1} and 7.7 kW kg^{−1}, respectively.⁵⁰ Naresh, Nibagani, *et al.* presented advanced AZIBs based on an α-MnO₂@SnO₂ nanocomposite, which exhibited the highest capacity value of 347 mAh g^{−1}.⁵¹ Polat, Bulent, *et al.* reported the ZnO@polypyrrole-P(VSANA) electrode for H₂O₂ detection and a symmetric SC configuration, with a 0.044 μM LOD value and a 171.8 F g^{−1} specific capacitance value.⁵² Zhao, Jingyuan, *et al.* reported the monometallic elemental Ti₃C₂T_x-MXene (Mo₂TiC₂T_x) demonstrating a capacitance of 398 F cm^{−3} with 82.75% capacitance maintenance and volumetric energy and power densities of 48.1 W h L^{−1} and 230.7 W L^{−1}, respectively.⁵³ Xiong, Chuanyin, *et al.* proposed 3D bifunctional CMT@N-rGO carbon microtubes hybridized with MnO₂ and PANI, having good supercapacitive and ORR response, presenting excellent rate performance during charge/discharge, 95% and 93% capacitance retention, and high energy efficiencies of 65–77% and 61–72% for both CMT@N-rGO@MnO₂ and CMT@N-rGO@PANI electrodes, respectively.⁵⁴ Mohammad, Akbar, *et al.* reported a CC-modified Al₂O₃-g-C₃N₄ nonenzymatic electrochemical sensor for H₂O₂ detection having an LOD value of 1.6 × 10^{−4} M and a sensitivity of 108 μA mM^{−1} cm^{−2}.⁵⁵

Continuing the usage of conventional transition-metal-based materials, the present study explores rhenium (Re) and its selenide derivatives as promising base materials owing to their distinctive 5d electronic configurations and multivalent oxidation states of Re⁴⁺ and Re⁶⁺. Unlike traditional 3d transition metals, Re possesses a broader d-band dispersion and a stronger spin–orbit coupling, which will promote rapid charge



transfer and enhance the intrinsic electrical conductivity of the electrode. Moreover, rhenium diselenide (ReSe_2) exhibits a distorted 1T lattice with abundant active edge sites and anisotropic in-plane electrical pathways, offering a unique advantage for facilitating multielectron redox reactions and improving the catalytic conversion of electroactive species. To further optimize electronic and ionic transport kinetics, the MIL-53 (Fe-MOF) host is selected as a structural backbone. It offers a high surface area, an adjustable pore architecture, and metal–ligand coordination-driven design flexibility. The porous MOF matrix provides uniform anchoring sites for ReSe_2 nanostructures, suppressing particle aggregation while establishing well-defined channels for electrolyte infiltration and charge migration. Such synergistic integration of ReSe_2 within the MOF is anticipated to yield a composite with enhanced structural integrity, accessible reaction interfaces, and accelerated redox kinetics during electrochemical operation. Along with this, we have incorporated graphitic carbon nitride ($\text{g-C}_3\text{N}_4$) as a dopant. This heteroatom dopant is introduced to modulate the electronic structure and generate interfacial defects that facilitate charge delocalization between the ReSe_2 and the MOF matrix. The dopant $\text{g-C}_3\text{N}_4$ atoms act as electron donors, enriching the local charge density and creating polar sites due to their intrinsic physicochemical properties. This plays a significant role in enhancing the electrochemical response to improve the adsorption and catalytic transformation of electroactive intermediates. The 2D graphitic-like structure of $\text{g-C}_3\text{N}_4$ provides a high surface area, abundant active sites, and a π -conjugated network that promotes efficient electron transport. The nitrogen-rich framework of $\text{g-C}_3\text{N}_4$ offers multiple lone-pair electrons, facilitating strong coordination and charge transfer interactions with transition metal species such as Re and Fe within ReSe_2 and MIL-53, respectively. This leads to accelerated redox kinetics and improved electrical conductivity across the hybrid interface. $\text{g-C}_3\text{N}_4$ also serves as a structural buffer and a conductive matrix, preventing the agglomeration of ReSe_2 nanosheets and MIL-53 particles while maintaining mechanical integrity during repeated charge–discharge cycles. This novel, tailored, and electronically modulated ReMIL-CN composite material is expected to establish a rational framework for developing a multifunctional electrode architecture that unites high electrical conductivity, abundant active sites, and robust structural stability, ultimately advancing the performance of electrochemical energy storage and conversion devices.⁵⁶

In this work, we aimed to elucidate the mechanism and performance of supercapacitors (BHSCs) and H_2O_2 sensors by investigating the structure and morphology of the synthesized ReMIL-CN nanocomposite material and its electrochemical performance.

2 Experimentation

2.1 Materials and methods

The following analytical-grade reagents and chemicals were used in the synthesis process of the composite materials. The rhenium (Re) and selenium (Se) powder precursors were sourced from Sigma-Aldrich; terephthalic acid, KOH, and

activated carbon (AC) were bought from Molychem; and iron(III) nitrate nonahydrate ($\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$) was obtained from Alfa Aesar. Hydrazine hydrate ($\text{N}_2\text{H}_4 \cdot \text{H}_2\text{O}$), *N,N*-dimethylformamide, urea ($\text{CO}(\text{NH}_2)_2$), ethanol, and DI water were utilized as solvents.

The ReMIL-CN composite material was synthesized *via* a sequential hydrothermal and solvothermal approach. $\text{g-C}_3\text{N}_4$ was synthesized *via* the green hydrothermal method using the low-cost, ecofriendly nitrogen-rich urea precursor. For this, a 1 M precursor solution of urea in DI water was prepared by dissolving and stirring to ensure complete dissolution and homogeneity. By heating in the autoclave at 170 °C for 8 hours, the polymerized $\text{g-C}_3\text{N}_4$ was formed. The $\text{g-C}_3\text{N}_4$ product was precipitated after cooling, washed, dried, ground, and preserved for use.

To synthesize ReSe_2 , a stoichiometric mixture of Re and Se powders was mixed in a 1 : 2 molar ratio and dispersed in 40 mL of deionized water under vigorous magnetic stirring. A small amount of hydrazine hydrate ($\text{N}_2\text{H}_4 \cdot \text{H}_2\text{O}$) was added as a reducing agent, which enhanced the solubility of materials and facilitated the reaction kinetics. The resulting suspension was autoclaved at 190 °C for 12 hours. The resultant ReSe_2 precipitated product was naturally cooled, washed several times to remove unreacted materials, and dried in a vacuum oven at 60 °C for 12 hours.

For MIL-53 (Fe-MOF) synthesis, a 50 mL 1 M solution containing $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$, terephthalic acid ($\text{C}_8\text{H}_6\text{O}_4$), DMF, and DI water was prepared. $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ was the metal source, and terephthalic acid ($\text{C}_8\text{H}_6\text{O}_4$) was the organic linker. It was subjected to continuous stirring to form a homogeneous solution. The prepared solution was poured into a 50 mL Teflon-lined stainless-steel autoclave, sealed tightly, and subjected to a hydrothermal process under heating at 150 °C for 12 hours. After natural cooling to room temperature, the resulting reddish-brown precipitate was collected by centrifugation. It was then thoroughly washed to remove unreacted precursors and residual solvents and dried in a vacuum oven at 60 °C for 12 hours.

To synthesize the ReSe_2 @MIL-53@ $\text{g-C}_3\text{N}_4$ hybrid composite, equimolar aqueous dispersions of the presynthesized ReSe_2 , Fe-based MIL-53, and $\text{g-C}_3\text{N}_4$ were prepared. These components were then mixed in a 2 : 1 : 1 ratio under continuous magnetic stirring for 30 minutes, ensuring uniform dispersion and effective interaction among all constituents. The resulting homogeneous suspension was subsequently transferred to a Teflon-lined stainless-steel autoclave and subjected to hydrothermal treatment at 180 °C for 12 hours, which facilitated the intimate interfacial coupling of ReSe_2 with MIL-53 and $\text{g-C}_3\text{N}_4$, leading to the formation of a well-integrated ternary ReMIL-CN composite structure. The resultant product was processed through subsequent steps into a finely ground powder for further use.

Fig. 1 comprehensively describes the novel synergistic integration in the ternary composite material. It highlights the structural morphologies involved, synthesis steps, and resulting hybrid structure that is supposed and justified to own the unique combined characteristics of all constituent materials. The



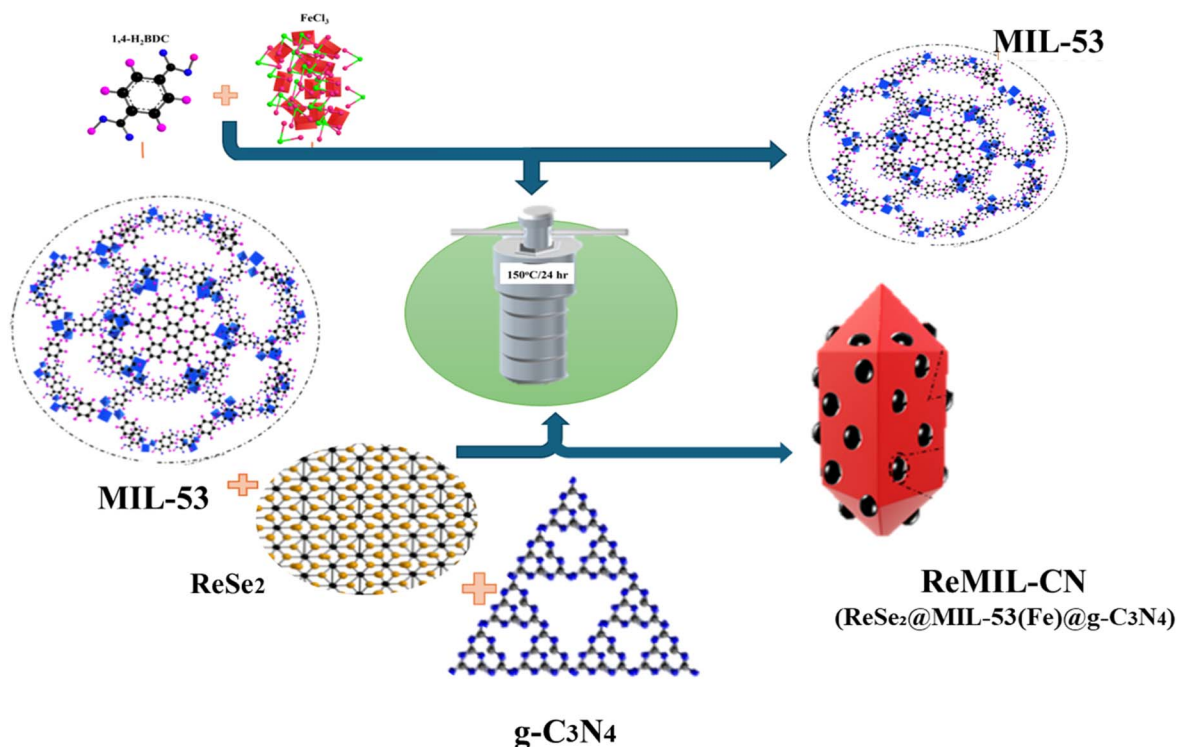


Fig. 1 Schematic showing the integration of ReSe₂, iron-MOF (MIL-53), and g-C₃N₄ into the hybrid ternary composite material ReMIL-CN.

resultant composite material is represented by a red octahedron with an embedded Fe-MOF (framework) and ReSe₂ nanoparticles. The structural components, including the porous Fe-MOF and the ReSe₂ layer, are also exhibited in Fig. 1.

The morphology and surface texture of the synthesized materials were examined using scanning electron microscopy (SEM) (JEOL JSM-6510LV) operated at an accelerating voltage of 10–15 kV. SEM energy-dispersive X-ray spectroscopy (EDX) was employed to determine the elemental composition and distribution, confirming the successful incorporation of the constituent elements. The crystalline structure of the samples was analyzed by X-ray diffraction (XRD) using a Cu-K α radiation source with a wavelength of $\lambda = 1.5406 \text{ \AA}$, operating at 40 kV and 30 mA. The obtained diffraction peaks were matched with the standard JCPDS data to identify the corresponding crystalline phases. The specific surface area and pore size distribution were determined from N₂ adsorption–desorption isotherms using a BET analyzer (Micromeritics ASAP 2460). The Brunauer–Emmett–Teller (BET) method was applied to calculate the surface area, while the Barrett–Joyner–Halenda (BJH) model was used to determine pore size and pore volume.

Electrochemical measurements were carried out using a standard three-electrode system connected to an electrochemical workstation (Corrtest). During electrochemical characterization, the prepared electrode served as the working electrode, a platinum wire as the counter electrode, and Hg/HgO as the reference electrode. All measurements were conducted in a 1 M KOH aqueous electrolyte solution at room temperature. Cyclic voltammetry (CV), galvanostatic charge-discharge (GCD), electrochemical impedance spectroscopy

(EIS), and Dunn's model tests were performed to evaluate the redox behavior, charge storage mechanism, specific capacity, cyclic stability, internal resistance and charge transfer dynamics, and energy and power densities.

After the parametric study of the fabricated electrodes, the study was extended to a real BHSC hybrid device and an electrochemical H₂O₂ sensing system. In the BHSC device, the active material positive electrode and activated carbon (AC) negative electrode were asymmetrically configured. The mass balance between the positive and negative electrodes was optimized according to the following charge balance relation to ensure the stable and efficient charge storage performance of the device:

$$m^+ C^+ V^+ = M^- C^- V^- \quad (1)$$

For H₂O₂ sensing evaluation, the three-electrode electrochemical setup was employed, comprising the active material-based modified electrode as the WE, a platinum (Pt) wire as the CE, and Hg/HgO as the RHE.

2.2 XRD and scanning electron microscopy (SEM)

The obtained XRD pattern confirmed the successful incorporation of all three components without impurity phases. The diffraction pattern in Fig. 2(a) reveals the distinct reflections for individual components g-C₃N₄, MIL-53(Fe), and ReSe₂ and the hybrid ternary compound. The obtained XRD peaks corresponding to each constituent, without any significant impurity signals, confirmed the successful synthesis of the ternary hybrid ReMIL-CN composite material. For g-C₃N₄, the characteristic



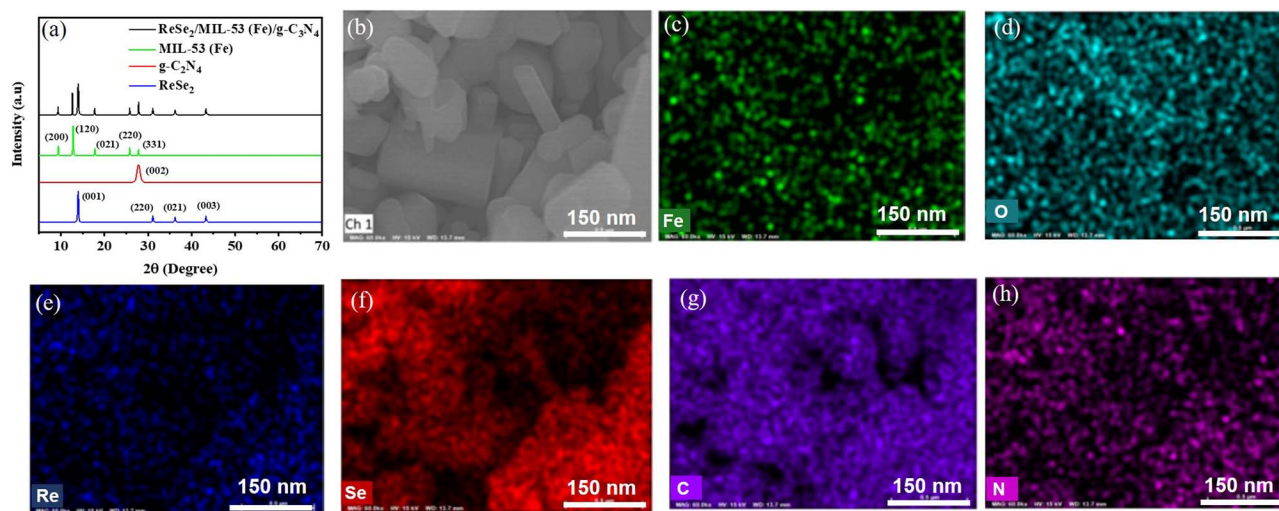


Fig. 2 (a) XRD characterization, (b) SEM image, and (c–h) SEM-EDX analysis.

peaks were observed at $2\theta = 12.8^\circ$ and 27.6° , analogous to the (100) and (002) hkl planes, respectively. These indexed peaks are the characteristic fingerprints of $g\text{-C}_3\text{N}_4$, indicating the in-plane tri-s-triazine repeated structural units and interlayer stacked conjugated aromatic planes. This corresponds to the periodic arrangement and structural repetition within the 2D layers of $g\text{-C}_3\text{N}_4$ (JCPDS no. 87-1526).⁵⁷ MIL-53(Fe) exhibited characteristic sharp peaks at $2\theta = 9.2^\circ$, 12.0° , 17.1° , 18.5° , and 25.6° , indexed to the (200), (120), (220), (021), and (331) planes, respectively. The obtained peaks at 2θ values of 9.2° , 12.6° , and 17.6° corresponded to the hydrated crystalline form of MIL-53(Fe) and match well with the previously reported results for Fe-based MOFs. For ReSe₂, the sharp reflections found at 2θ values of 13.6° , 31.2° , 38.7° , and 44.0° corresponded to the (001), (220), (021), and (003) planes, respectively, confirming the triclinic phase structure of ReSe₂ with a dominant growth orientation along the (001) plane (c -axis alignment). These peak intensities and positions matched well with JCPDS file no. 04-007-1113.⁵⁸ The obtained XRD pattern of the present ReMIL-CN composite material demonstrated a combined diffraction pattern that retained all major reflections from its constituent materials. The sharpness of the peaks for the hybrid composite also confirmed the purity and ordered atomic arrangement within the structure. These findings confirmed the successful synthesis of the well-crystalline ReMIL-CN hybrid nanocomposite with the synergistic structural integration of ReSe₂, MIL-53(Fe), and $g\text{-C}_3\text{N}_4$.

The XRD results were used to ascertain the degree of crystallinity, lattice parameter, phase angle, and phase composition of the presented ReMIL-CN nanocomposites. The average crystallite size (D) of the composite was measured from the intensity and width of these peaks using the Debye–Scherrer equation:

$$D = \frac{K\lambda}{\beta \cos \theta} \quad (2)$$

Here, “ D ” represents the crystallite size, “ K ” is the Scherrer constant (typically 0.9), “ λ ” denotes the X-ray wavelength (1.5406 Å), “ β ” is the full width at half-maximum (FWHM), expressed in radians, and “ θ ” is the Bragg diffraction angle.

The synthesized ReMIL-CN composite material's average crystallite size was measured perpendicular to the (002) planes at $2\theta = 27.6^\circ$, which was 27.27 nm, and the FWHM was 0.3°. This demonstrates the nano-sized dimensions of the stacked composite, corresponding to the vertical crystalline coherence length.

The SEM characteristic image of the presented ReMIL-CN nanocomposite, shown in Fig. 2(b), provided valuable insights into the morphological architecture and structural integrity of the synthesized nanocomposite. It displayed a highly porous, flake-like morphology with densely packed interconnected nanosheets and agglomerated irregular-shaped particulates, forming a hierarchical structure. The nano-scale features with dimensions well below 150 nm confirmed the formation of a nanostructured hybrid composite. The sheet-like layered structure in the SEM image indicated the presence of exfoliated $g\text{-C}_3\text{N}_4$ layers in the composite. These nanosheets appeared to be thin and crumpled, providing a large surface area and abundant edge sites for electrochemical interactions. The irregularly embedded granular particles within these layers were likely due to the MIL-53(Fe) and ReSe₂ constituents. The granular features suggested that the MIL-53(Fe) MOF had been successfully incorporated into the composite and retained its porous structure within the hybrid matrix. The existence of flake like structure is typically exhibited by the ReSe₂ and integrated into the porous framework, forming strong interfacial contact with the $g\text{-C}_3\text{N}_4$ nanosheets and the MOF matrix. The uniform and loosely stacked layered structure indicated the mesoporous nature of the presented synthesized hybrid ternary composite, leading to enhanced ion diffusion and an enhanced available surface area for electrochemical reactions. The mesoporous architecture of the present composite, without isolated and



large particles, as shown in the SEM image, suggested good compatibility among its components and the successful hybridization of the material. This well-integrated structure indicated the potential robustness and stability of the composite material during practical applications, including biosensing, photocatalysis, and supercapacitor electrodes.

The images in Fig. 2(c–h) display the results of SEM-EDX spectroscopy, providing the qualitative elemental mapping of the presented synthesized hybrid composite material. This energy-dispersive X-ray coupled SEM analysis provided the detailed spatial distribution and relative abundance of the constituent elements across the sample. The EDX elemental mapping confirmed the successful incorporation and even dispersion of the constituent elements across the composite's surface, demonstrating the successful synthesis and potential synergy among components. The uniform distribution of iron (Fe) in Fig. 2(c) and rhenium (Re) in Fig. 2(d) confirmed their successful codeposition within the material matrix. The widespread and dense distribution of selenium (Se) is shown in Fig. 2(d), indicating its potential role as a major active site-contributing component of the material. The uniformly dispersed carbon (C) and oxygen (O) typically originated from the carbon-based ($g\text{-C}_3\text{N}_4$) support structure and possible surface oxidation. Elemental nitrogen (N), as shown in Fig. 2(h), was less densely distributed but was still present across the

sample, indicating potential N doping, which can enhance the conductivity and surface reactivity of the composite material. This elemental homogeneity demonstrated the consistent composition of the synthesized material, making it critical for achieving reliable responses in electrocatalysis-based sensors and energy storage devices.

2.3 Electrochemical studies: CV, GCD, and EIS

The electrochemical performance of the fabricated electrodes was systematically evaluated by cyclic voltammetry (CV) and galvanostatic charge–discharge (GCD) analyses. A 1 M KOH-DI water electrolyte solution was used for electrochemical measurements because of its efficient ionic conductivity, good charge transfer ability, and efficient operational suitability with transition metal-based materials.

The CV plots shown in Fig. 3(a–c) display the current response of the ReSe_2 , $\text{MIL-53(Fe)}@ \text{ReSe}_2$, and hybrid compound electrodes at varying scanning potentials from 3 mV s^{-1} to 50 mV s^{-1} versus RHE. Each electrode exhibited well-defined redox peaks, corresponding to the occurrence of redox reactions within these electrode materials. The pronounced Faradaic characteristic peaks obtained in CV plots reflected the battery-like behavior of these electrodes. Among the results of the three samples, the ternary composite ReMIL-CN electrode demonstrated more pronounced and symmetric redox peaks,

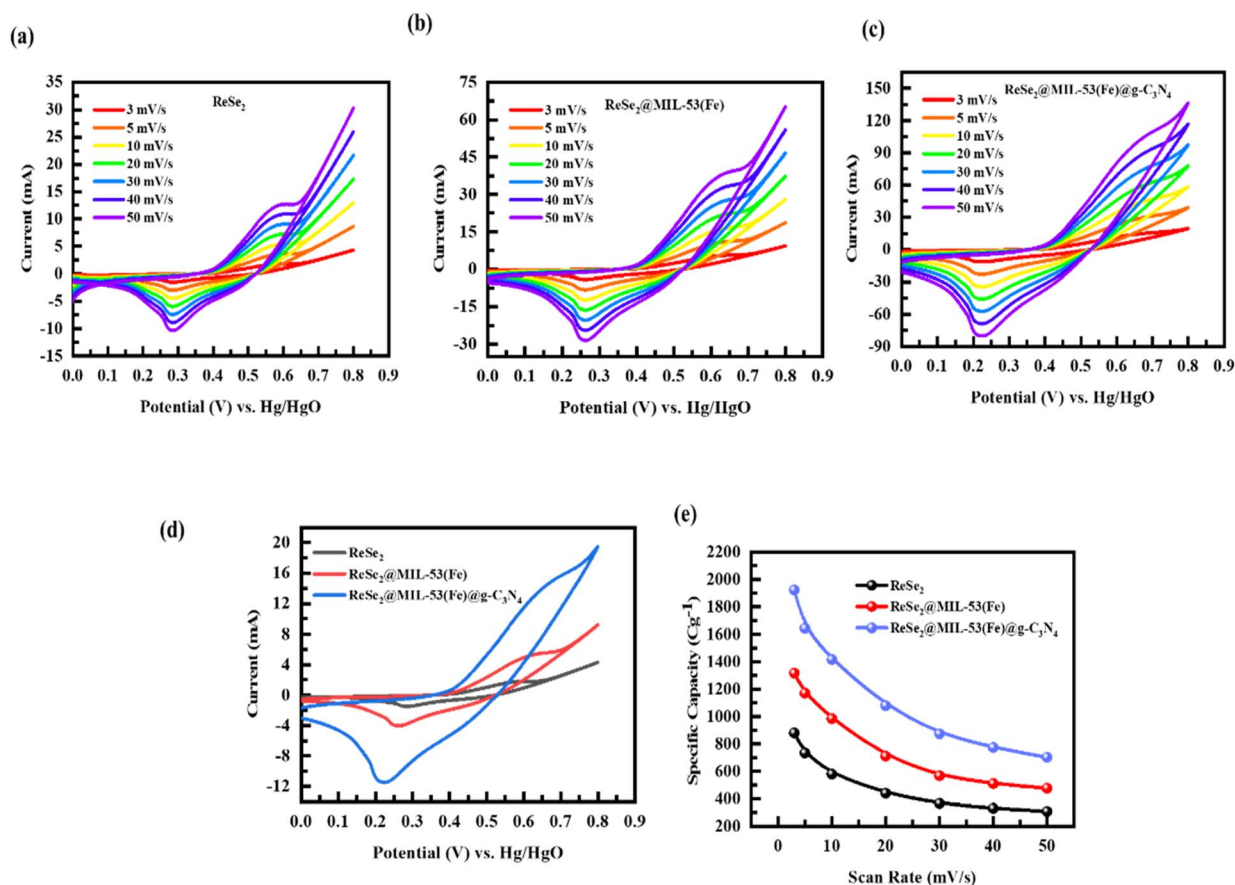


Fig. 3 CV curves of (a) ReSe_2 , (b) $\text{MIL-53(Fe)}@ \text{ReSe}_2$, and (c) the ReMIL-CN composite for the scan rate range of $3\text{--}50 \text{ mV s}^{-1}$, (d) comparative CV performance of the three electrodes at 50 mV s^{-1} , and (e) specific capacity versus scan rate for the three electrodes, calculated from CV results.



indicating the enhanced electrochemical reversibility and redox kinetics of this electrode. These peaks originated from the synergistic redox transitions of iron ($\text{Fe}^{3+}/\text{Fe}^{2+}$) in the MIL-53(Fe) framework and $\text{Re}^{4+}/\text{Re}^{3+}$ in ReSe_2 , accompanied by reversible ion intercalation/deintercalation processes within the hybrid matrix.⁵⁹ The strong current response and nearly symmetric redox peaks indicated high reversibility and fast ion diffusion kinetics. This suggested that both surface and bulk faradaic reactions actively contributed to charge storage. The symmetric CV curves indicated balanced charge/discharge cyclic processes and minimal internal resistance, leading to a high electrochemical charge storage response obtained for the presented synthesized novel composite-based electrode, with stable charge storage mechanisms and durability during repeated cycles. The rise in the current with an increase in the voltage demonstrated an enhanced charge storage capacity and conductivity at high voltage values. The narrow gaps between forward and reverse branches demonstrated a minimal overpotential, ensuring efficient energy conversion in the electrode-based system. Based on the comparison of the CV curves of three composite materials' electrodes, the ReMIL-CN composite electrode showed a more pronounced faradaic mechanism, especially under higher input conditions, making it a promising candidate for energy storage applications.

The specific capacity (Q_s) versus scan rate values presented in Fig. 3(e) were measured from the CV characteristic results for

the ReSe_2 , MIL-53(Fe)@ ReSe_2 , and ReMIL-CN hybrid material electrodes using the following relation:

$$Q_s = \frac{1}{mv} \int_{v_i}^{v_f} I \times VdV \quad (3)$$

Here, Q_s is the specific capacity (C g^{-1}), m is the mass of the active material (g), v is the scan rate (V s^{-1}), and I is the current (A). The integration of the current over the potential range provided the total charge, which was normalized by the mass and scan rate, yielding Q_s . The values of the specific capacity decreased with increasing scan rate for all three electrodes due to limited ion diffusion at higher scan rates, which restricts the full utilization of active sites. Among the three tested materials, the hybrid composite material demonstrated the highest specific capacity consistently across all scan rates, reaching the highest value of 1925.5 C g^{-1} at 5 mV s^{-1} and maintaining the minimum value of 800 C g^{-1} at 50 mV s^{-1} . Compared to the lower capacities achieved by MIL-53(Fe)@ ReSe_2 and ReSe_2 , the hybrid material electrode demonstrated a superior performance, likely due to the synergistic effects facilitated by the composite architecture of ReSe_2 , Fe-MOF (MIL-53) and $\text{g-C}_3\text{N}_4$, suggesting that it is a favorable material for high-performance electrochemical energy storage systems. In addition to CV, GCD tests were also conducted for the ReSe_2 , MIL-53(Fe)@ ReSe_2 , and hybrid compound-based electrodes at varying current densities from 1 A g^{-1} to 2.2 A g^{-1} , leading to a more direct assessment of the electrochemical performance of these electrodes. The GCD analysis illustrated in Fig. 4(a-c) provided

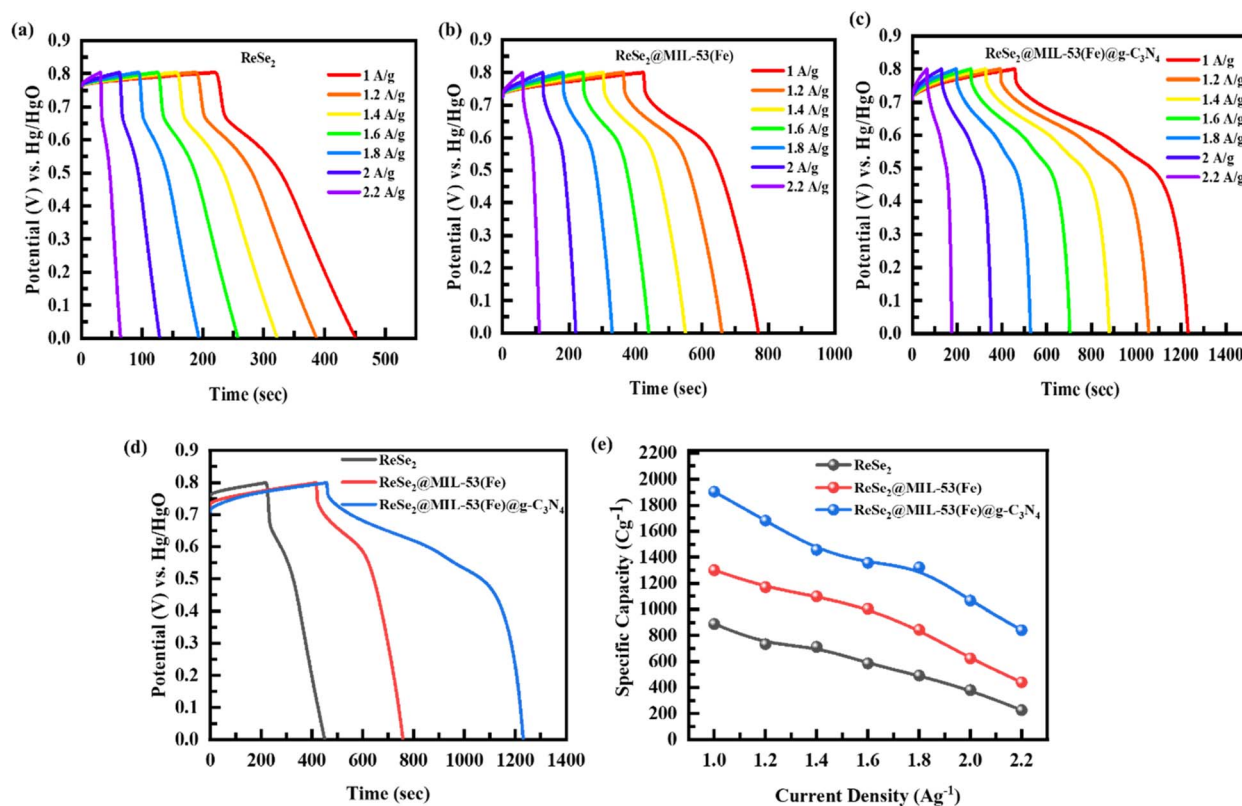


Fig. 4 GCD characteristics curves of (a) ReSe_2 , (b) MIL-53(Fe)@ ReSe_2 , and (c) the ternary composite ReMIL-CN over the 1 – 2.2 A g^{-1} input current range, (d) comparative GCD performance of the three nanocomposites at 50 mV s^{-1} , and (e) their measured specific capacity values.



symmetric charge–discharge profiles with distinct plateaus, further confirming the faradaic-dominated mechanism. The minimal potential difference between the charge and discharge plateaus for the ReMIL-CN electrode reflected the low polarization and excellent electrochemical kinetics of the hybrid electrode. The linear regions of GCD profiles demonstrated the good electrochemical activity and conductivity of the electrodes, and the relatively flat regions with linear slopes signified quick electron mobility throughout their devices. A gradual reduction in the discharge time of all three electrodes was observed as the current density increased, which is a typical behavior due to ion diffusion and kinetic limitations at higher current rates.

The GCD graphs suggested that the ReMIL-CN composite's electrode outperformed the individual components' electrodes. This electrode retained a relatively better capacity even at elevated current densities, suggesting the excellent rate capability and structural stability of this electrode-based system. The longest discharge times of this electrode indicated its highest specific capacity and superior electrochemical performance. It also corresponded to the stable ion storage response of the presented electrode-based device.

Fig. 4(d) presents the comparative GCD profiles of the three electrodes at 1 A g^{-1} , highlighting the performance superiority of the ReMIL-CN composite electrode over pristine ReSe_2 and $\text{MIL-53(Fe)}@ \text{ReSe}_2$ electrodes. This enhanced performance arose from the synergistic combination within the composite electrode, making it a strong candidate for upgrading ionic accessibility, capacity, electrochemical stability, and discharge time stability for energy storage applications.

The electrochemical charge storage capability and rate performance of all three composite-based electrodes were

assessed by calculating their specific capacity (Q_s) versus scan rate from their GCD results using the following equation:

$$Q_s = \frac{I \times \Delta t}{m} \quad (4)$$

Here, I is the applied current, Δt is the discharge time, and m is the active mass of the material. The smooth and linear reduction in the specific capacity values with an increase in the current also showcased the material's capability to manage rapid ion movement without considerable capacity loss. The highest obtained values of the specific capacity were 1912 C g^{-1} for the ReMIL-CN electrode, 1206.5 C g^{-1} for $\text{MIL-53(Fe)}@ \text{ReSe}_2$, and 858 C g^{-1} for the ReSe_2 electrode. These measured specific capacities from CV and GCD were in proximity, and the slight variation (<1%) was attributed to differences in the charge–discharge kinetics and potential window integration during measurement. This close agreement confirmed the reliability of the electrochemical evaluation and highlighted the excellent charge storage ability of the ReMIL-CN composite's electrode.

The anodic and cathodic peak current versus square root of scan rate variations are depicted in Fig. 5(a–c) for the ReSe_2 , $\text{MIL-53(Fe)}@ \text{ReSe}_2$, and the ReMIL-CN sample materials, respectively, providing insights into their electrochemical performance. These plots also showed the linear relationship for the ion insertion and extraction processes. The plot in Fig. 5(a) for ReSe_2 exhibited its linear trends with relatively lower slopes. It indicated not only diffusion-controlled kinetics but also limited ion diffusion rates and conductivity, which constrained its charge storage capacity. The ReMIL-CN compound-based electrode resulted in a steeper concentration gradient due to its more pronounced CV peak currents and higher charge kinetic values for each input scan rate value.

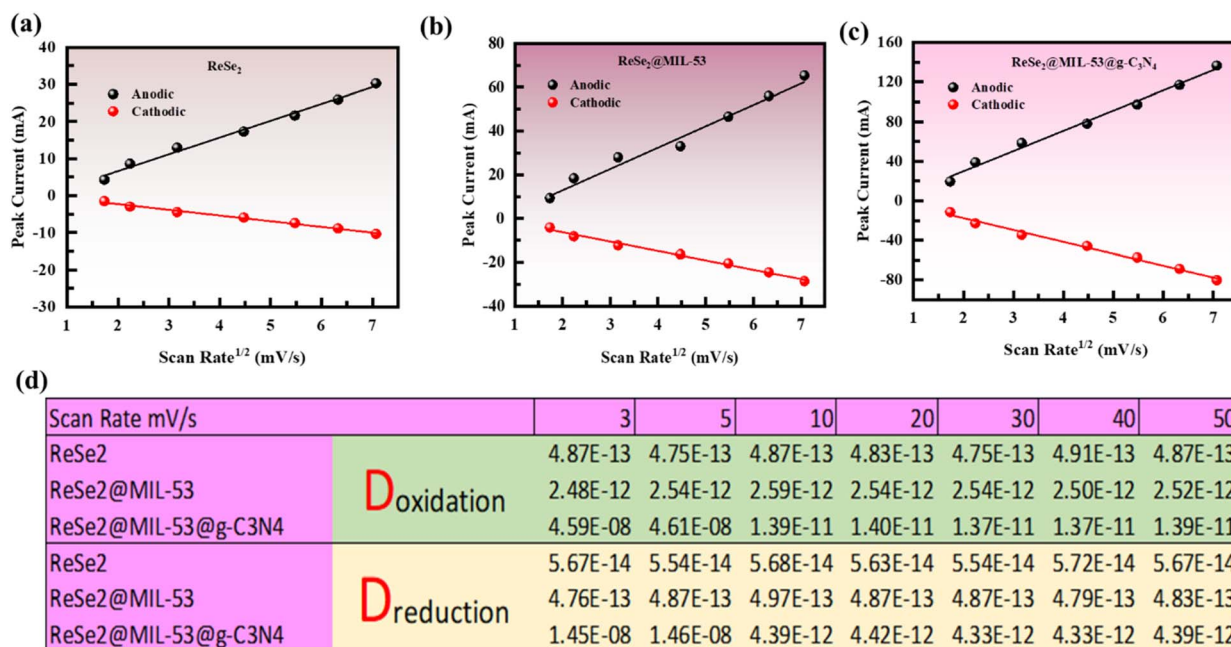


Fig. 5 Anodic and cathodic peak currents versus scan rate for the (a) ReSe_2 , (b) $\text{MIL-53(Fe)}@ \text{ReSe}_2$, and (c) ReMIL-CN composite material and (d) oxidation and reduction diffusion coefficients of the three materials.



These anodic and cathodic peak current responses were analyzed to estimate the apparent ion transport kinetics, which revealed that the respective electrode facilitated fast and stable redox reactions during repeated cycling. Although the diffusion process could not be strictly described using classical solution-phase models, the relative changes in peak current intensities provided valuable insights into the electrochemical kinetics and ion diffusion efficiency. The obtained higher apparent anodic and cathodic responses corresponded to an intrinsic redox mechanism and suggested improved ionic mobility and facile electron transfer pathways, consistent with the synergistic effect of the ReMIL-CN framework and the conductive morphology. These findings indicated that the ReMIL-CN electrode exhibited a balanced combination of surface-controlled and diffusion-assisted charge storage, contributing to its superior electrochemical performance.

The highest obtained values of the diffusion coefficient “ D ” for both oxidation and reduction reactions were $D_{\text{oxidation}} = 4.87 \times 10^{-13} \text{ m}^2 \text{ s}^{-1}$, $2.48 \times 10^{-12} \text{ m}^2 \text{ s}^{-1}$, and $4.59 \times 10^{-8} \text{ m}^2 \text{ s}^{-1}$ and $D_{\text{reduction}} = 5.67 \times 10^{-14} \text{ m}^2 \text{ s}^{-1}$, $4.76 \times 10^{-13} \text{ m}^2 \text{ s}^{-1}$, and $1.45 \times 10^{-8} \text{ m}^2 \text{ s}^{-1}$ for the ReSe_2 , $\text{MIL-53(Fe)}@ \text{ReSe}_2$, and ReMIL-CN hybrid material, respectively. The highest obtained diffusion coefficient “ D ” for the ReMIL-CN compound confirmed the quicker kinetics of electroactive species towards

this electrode. This rapid diffusion of charges to the electrode surface increased the resultant currents. A high value of D is also directly related to the sharpness of the redox peaks in CV plots, supporting surface suitability, a porous and rough surface nature, and a large effective surface area for redox reactions, leading to the tailoring of the reversibility of the electrode.

2.4 Electrochemical impedance spectroscopy (EIS)

The EIS analysis investigates the charge transfer dynamics of the prepared material-based electrode systems. The EIS results obtained for ReSe_2 and the ReMIL-CN composite material are presented in Fig. 6(a), which provides a detailed insight into the charge transfer resistive and capacitive interfacial properties of ReSe_2 and the ReMIL-CN hybrid composite material. In the high-frequency region of the Nyquist plot, a semicircle is observed, indicating the charge transfer resistance (R_{ct}) at the electrode/electrolyte interface. At high frequencies, the ReSe_2 exhibits a higher R_{ct} value, as evident by its EIS arc with a larger semicircle, indicating the slower charge transfer kinetics of ReSe_2 . In contrast, the ternary nanocomposite ReMIL-CN shows a smaller semicircle, reflecting a significantly reduced R_{ct} and improved electrochemical performance resulting from the enhanced conductivity and synergistic interactions within the ReMIL-CN hybrid material. These results suggest improved

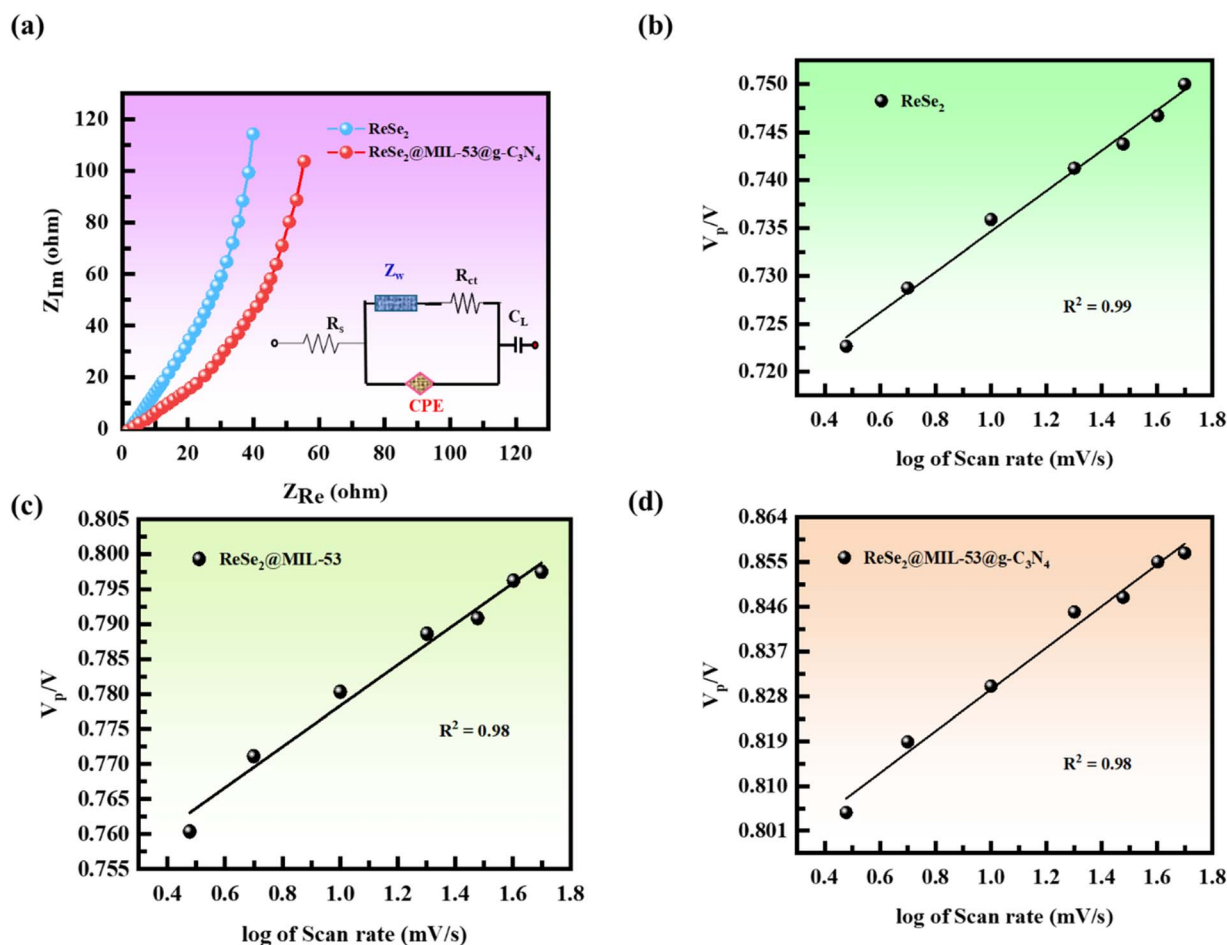


Fig. 6 (a) EIS measurement of ReSe_2 and $\text{MIL-53(Fe)}@ \text{ReSe}_2@ \text{g-C}_3\text{N}_4$ composite (b–d) peak voltage as a function of \log of scan rate for ReSe_2 , $\text{MIL-53(Fe)}@ \text{ReSe}_2$, and ReMIL-CN composite.



electron transport and catalytic activity in the composite material, making it more favorable for efficient and fast energy storage applications. The EIS equivalent circuit model includes the series resistance (R_s) corresponding to the solution resistance. The semicircular region observed in the Nyquist plot corresponds to the charge transfer resistance (R_{ct}). The CPE demonstrates the constant phase element that accounts for the nonideal capacitive behavior at the electrode/electrolyte interface. Additionally, the Warburg impedance (Z_w) is associated with diffusion-controlled ion transport processes. The inclusion of these components assists in adjusting the overall electrochemical role of the ReMIL-CN nanocomposite's electrode and its ion diffusion, particularly at higher current densities.

To further elaborate on the charge storage mechanism, the correlational potential *versus* scan rate variation was studied, as shown in Fig. 6(b–d), for the three composites ReSe₂, MIL-53(Fe)@ReSe₂, and ReMIL-CN, corresponding to the peak currents in their respective CV analysis. The relative peak potential (V_p/V) of the output current *versus* scan rate variations for ReSe₂, MIL-53(Fe)@ReSe₂, and their ReMIL-CN nanocomposite electrodes, as shown in Fig. 6(d), reveals the strong linear relationships for the three electrodes, indicating the diffusion-controlled charge kinetics and stable electrochemical behavior of these electrodes. The pristine ReSe₂ electrode exhibits a correlation value of $R^2 = 0.99$, signifying an unhindered diffusion mechanism, while MIL-53(Fe)@ReSe₂ and the ReMIL-CN composite exhibit R^2 values of 0.98, suggesting similar diffusion-controlled processes. The variations in R^2 values for the hybrid composites can be attributed to structural and electronic enhancements due to the addition of Fe-MOF (MIL-53) and g-C₃N₄,

which improve the performance, charge storage, and stability of these electrodes.

2.5 BET (nitrogen adsorption–desorption isotherm) analysis

The Brunauer–Emmett–Teller (BET) characteristic analysis was carried out *via* N₂-physisorption measurements using the Quantachrome Autosorb analyzing system. The obtained adsorption/desorption isotherms are presented in Fig. 7(a–d) for the ReSe₂, MIL-53(Fe), g-C₃N₄, and hybrid material electrodes, respectively. The material's porosity, including specific pore dimensions and their distribution, was measured from these isotherms.

All four samples exhibit type IV isotherms with noticeable hysteresis loops. These linearly rising hysteresis loops, observed within the relative pressure (P/P_0) range of 0.4–0.9, indicate capillary condensation within mesopores, confirming the mesostructured nature of the synthesized samples. The BET isotherm of ReSe₂, presented in Fig. 7(a), exhibits a type IV curve accompanied by a distinct H3-type hysteresis loop. This behavior is typical of mesoporous materials composed of nonrigid plate-like particle aggregates that form slit-shaped open pores due to the layered structure of ReSe₂.⁶⁰ The obtained BET hysteresis for MIL-53(Fe) suggests gradual monolayer-multilayer adsorption with an increase in the relative pressure. This hysteresis loop indicates that MIL-53 (Fe) possesses a mesoporous structure with flexible framework behavior, highlighting the material's dynamic adsorption properties.

The BET hysteresis for g-C₃N₄ demonstrates a less steep uptake curve and lower total nitrogen adsorption, implying the

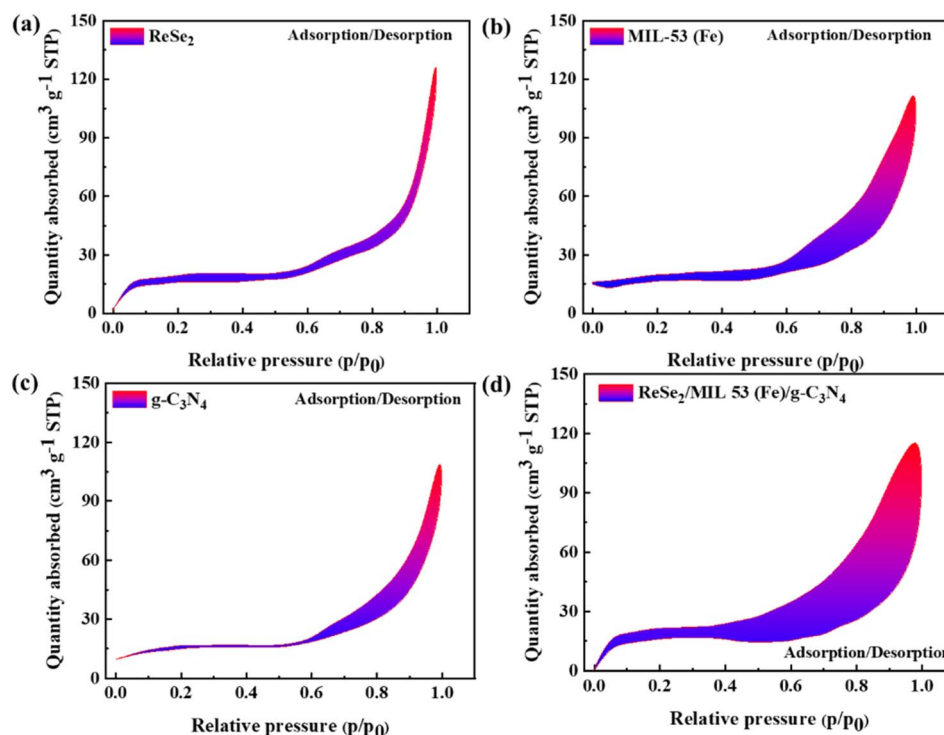


Fig. 7 BET analysis of the (a) ReSe₂, (b) MIL-53(Fe), (c) g-C₃N₄, and (d) ReMIL-CN composite electrodes.



relatively lower surface area and narrower pore distribution of this material. The BET results of the ReMIL-CN nanocomposite material, shown in Fig. 7(d), with a more pronounced hysteresis loop and a sharp increase in adsorption at higher relative pressures, demonstrate its highest N_2 uptake among all four materials. The observed trends reflect the superior adsorption performance of the ReMIL-CN hybrid composite, which is critical for specific applications in catalysis, gas storage, and separation. The total specific surface area and pore volume of the ReMIL-CN composite-based active electrode, estimated by the adsorbed/desorbed volume per gram ($\text{cm}^3 \text{g}^{-1}$) of N_2 versus relative vapor pressure (p/p_0), are $154 \text{ m}^2 \text{g}^{-1}$ and $0.022 \text{ cm}^3 \text{g}^{-1}$, respectively. The significant surface area and pore size of this ReMIL-CN composite material reflect its readiness for efficient ion transport, storage, and release during high-energy and -power applications.

In the ReMIL-CN hybrid composite material, the layered structure of ReSe_2 facilitates the enhancement of ion intercalation and deintercalation across varying current densities by allowing ions to enter and exit the material easily during the charge-discharge process. The smooth and dispersed metallic linker Fe-MOF increases the overall surface area and available active sites of the electrode material, allowing it to store large quantities of ions and provide abundant pathways for ion diffusion. The Fe-

MOF aids the device in maintaining a high capacity even at fast charging/discharging rates. The carbonaceous component ($\text{g-C}_3\text{N}_4$) introduces additional electronic pathways, leading to high electrical conductivity and fast ion transfer among the hybrid material's components. Overall, the electrochemical results affirm that the hierarchical structure of the ReMIL-CN composite's electrode offers substantial advantages for electrochemical energy storage applications by optimizing redox activity, structural stability, and electrical conductivity.

3 Real device applications

3.1 ReMIL-CN//AC hybrid device

The ReMIL-CN composite active electrode (AE) was further tested for real device ion storage applications in the standard two-electrode cell system with the second electrode of AC. The hybridization of the high specific energy and favorable cycling stability of the active electrode (AE) with the typically double-layered capacitive nature of the AC electrode leads to the benefits of both capacitive and faradaic processes in the hybrid system. The hybrid ReMIL-CN compound//AC electrode device is depicted in Fig. 8(a), with the WE being the active material ReMIL-CN. The two electrodes were immersed in a KOH electrolytic solution and were separated by a separator.

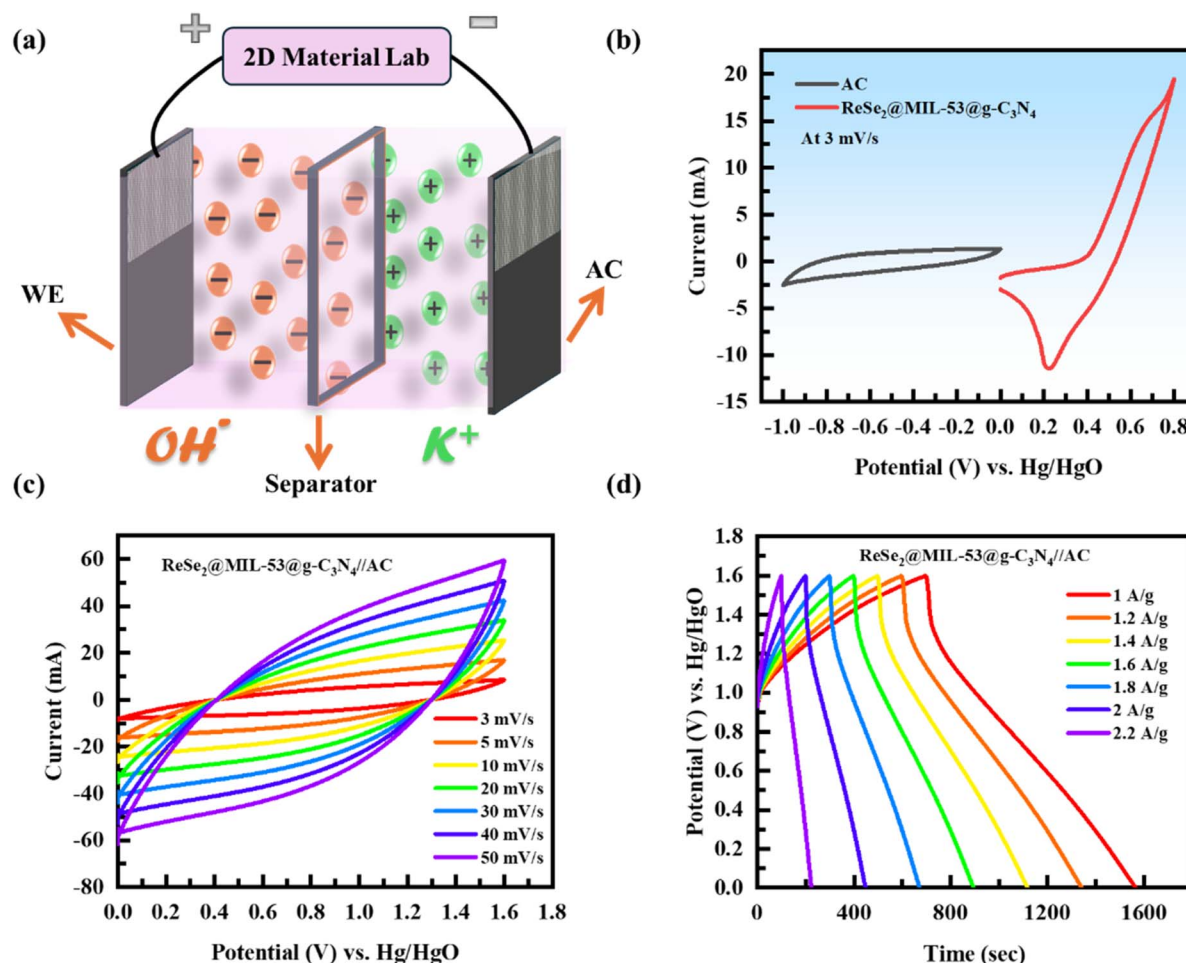


Fig. 8 (a) Real device depiction, (b) individual CV plots of the cathode and anode, and (c and d) hybrid device CV and GCD plots.



The CV results of ReMIL-CN composite and activated carbon (AC) electrodes are analysed using the three electrode systems as presented in Fig. 8(b). The results clearly indicate the capacitive response contribution of the AC electrode and faradaic response results presented by working electrode (WE). The counter electrode (AC) adds to the electrochemical capacitive performance by offering a highly stable and conductive surface area for substantial ion adsorption, thereby improving the overall energy density of the device. The CV and GCD plots in Fig. 8(c and d), respectively, provide a comprehensive understanding of the electrochemical behavior of the ReMIL-CN//AC hybrid device. The sharp and symmetric Faradaic redox peaks with quasi-rectangular-shaped CV curves and symmetric GCD profiles, even at higher scan rates, indicate rapid and reversible redox reactions dominated by surface-controlled charge storage rather than bulk ion diffusion. The excellent rate performance and minimal IR drop further confirm efficient electron/ion transport at the electrode–electrolyte interface.

Thus, the overall ReMIL-CN//AC device operates through a hybrid electrostatic–faradaic pseudocapacitive mechanism, ensuring high energy and power density with superior cycling stability. The area encapsulated by CV plots justifies its dominant charge storage mechanism and the storage capacity across

the input voltages. These transitions highlight the significance of electrode design optimization, preparation methods, input potential ranges, and configurations, which can improve the electrochemical performance. The GCD analysis presented in Fig. 8(d) affirms the suitability of the ReMIL-CN composite//AC hybrid device for practical ion/energy storage applications. The consistent, symmetric, and stable GCD plots of the devised system show efficient electron flow with a minimal voltage drop and low internal resistance during the charging/discharging process. The influence of the counter electrode in the fabricated hybrid system is evidenced by the consistent, repeatable shape of the curves across cycles for a sustainable and stable capacity role over long-term cycling. This stability is crucial for practical applications, where the device needs to operate reliably during long-term charge–discharge cycles. The combination of the unique properties of each component material in the ReMIL-CN hybrid composite results in high ion storage and efficient ion delivery. Furthermore, the charge storage-specific capacity of the designed ReMIL-CN composite//AC cell system, evaluated from both CV and GCD plots, is about 423 C g^{-1} .

The cycling stability, coulombic efficiency, and capacity retention of the ReMIL-CN composite//AC hybrid system were evaluated by subjecting it to 10 000 repeated electrochemical

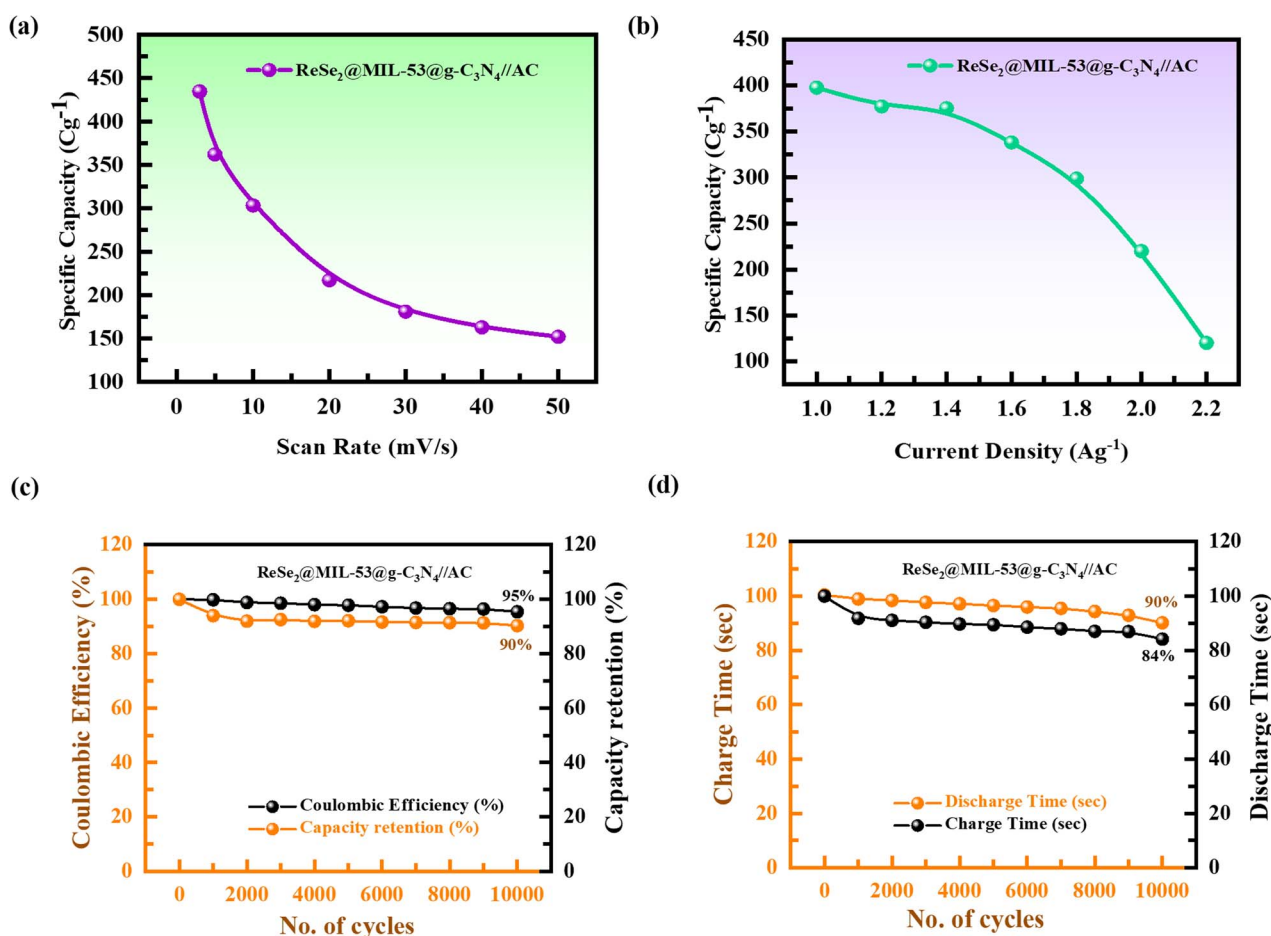


Fig. 9 (a and b) Specific capacity measurement results from CV, and GCD plots, (c and d) stability tests of the present hybrid device.



cycles under steady current conditions (Fig. 9). The presented fabricated device demonstrates an excellent 95% capacity retention, a 90% Columbic efficiency, and 90% and 84% charge–discharge cyclic retention stability, respectively. This suggests the structural and operational integrity of the ReMIL-CN composite//AC hybrid device over the prolonged operation of 10 000 cycles.

The b -value measurements of the devised system, as presented in Fig. 10(a), were employed to evaluate the charge storage behavior of the developed system. The b -values corresponding to the applied voltages of 0.6, 0.7, and 0.8 V were determined from the slope of the log–log plot of peak current ($\log i$) versus input scan rate ($\log \nu$). The slope $b = 0.59$ at 0.6 V suggests a predominantly diffusion-controlled process, which increases to $b = 0.68$ at 0.7 V, implying a transition towards a mixed contribution of surface-controlled capacitive and diffusion effects. Finally, at 0.8 V, the slope changes to $b = 0.7$, suggesting that surface-controlled processes become more prominent. The measured b -values indicate the hybrid charge storage behavior of the device, with diffusion-driven (battery-like) and surface-dominated capacitive characteristics. This behavior is desirable for tuning the hybrid (BHSC) response of the device. These observed b -values are indicative of the device's favorable ion kinetics, supporting rapid ion intercalation and deintercalation processes.

The study of the novel hybridized ReMIL-CN nanocomposite//AC system presented in this work highlights its outstanding rate capability at elevated voltages and current densities, along with high electrochemical storage response. The E_d and P_d values were measured from CV and GCD curves using the following eqn (5) and (6):

$$E_d = \frac{Q_s \times V}{2 \times 3.6} \quad (5)$$

$$P_d = \frac{E \times 3600}{\Delta t} \quad (6)$$

where Q_s represents the specific capacity, V is the operational voltage range, and t is the discharge time.

The highest values of the specific energy (E_d) and specific power (P_d) obtained for the ReMIL-CN//AC-cell system are 50 Wh kg^{-1} and 2060 W kg^{-1} , respectively. These values of E_d and P_d are comparable to or higher than those of many of the electrode materials documented in recent studies.

3.2 Dunn's model analysis

The charge storage behavior of the device was studied by separating the total current response into surface-controlled (capacitive) and diffusion-controlled contributions. Dunn's model plots in Fig. 11(a–c) present the relative contribution of these two storage mechanisms to the overall storage performance of the device with scan rate variations. At 3 mV s^{-1} , the diffusive contribution is close to 80%, while the capacitive contribution is around 20%, showing a strong contribution of the bulk (battery-type) behavior of the device. This is because at lower rates, the ions have more time to penetrate and accumulate within the electrode material. The capacitive contribution begins to increase with an increase in scan rates, and at 20 mV s^{-1} , the diffusive and capacitive parts reach nearly equal ratios. This reflects a balanced contribution of both mechanisms. At 50 mV s^{-1} , the capacitive contribution further increases, which corresponds to the dominance of surface redox processes, facilitating rapid ion transport and surface interactions at high input scan rates. The quantitative comparative analysis of surface-controlled and diffusion-controlled storage mechanisms, presented in Fig. 11(d), demonstrates a comprehensive analysis for the present system to achieve a balanced storage capacity. This Dunn's model analysis is valuable for adjusting our device's hybrid operational voltage range for fast ion intercalation and deintercalation applications that require high values of both energy and power densities.

The CV profiles of the as-configured ReSe_2 //AC and ReSe_2 @MIL-53//AC devices are shown in Fig. S1(a and d). These CV curves exhibit nearly rectangular shapes with slight redox humps, indicating a combination of electric double-layer capacitance and pseudocapacitive behavior. As the scan rate increases, the current response enhances proportionally

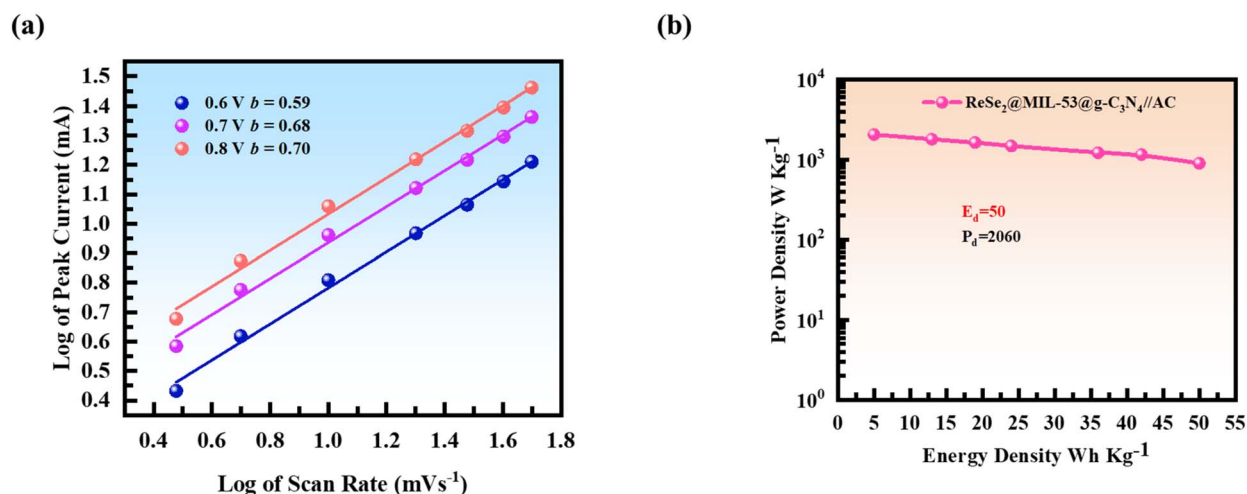


Fig. 10 (a) b -value measurement and (b) power density versus energy density Ragone plot.



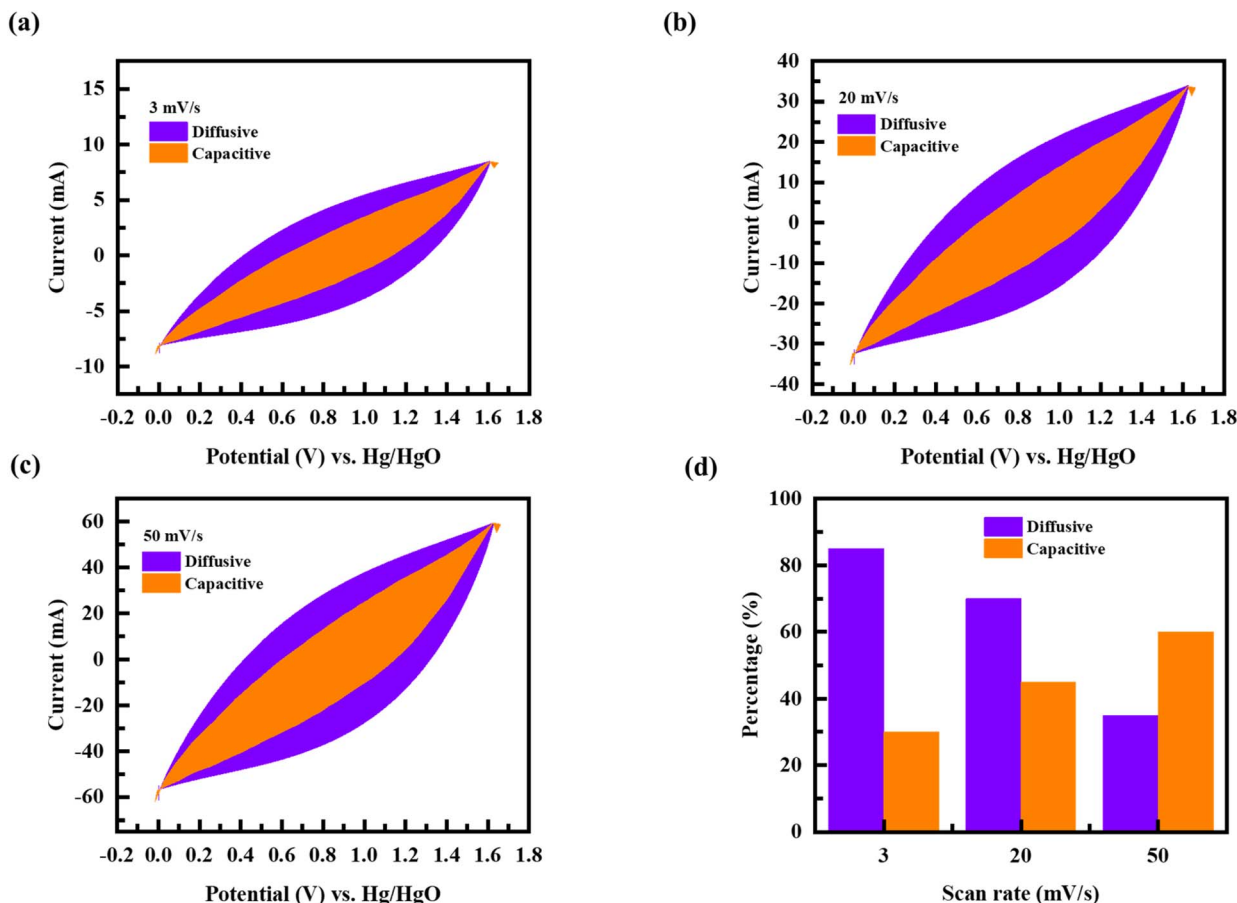


Fig. 11 (a–c) Results of Dunn's model and (d) comparison of the capacitive and diffusive response at 3, 20, and 50 mV s^{-1} .

without significant distortion, confirming excellent reversibility and rapid charge transfer capability. The stable shape of the CV profiles demonstrates good electrochemical stability and efficient ion diffusion within the electrode. According to the comparison plots of the three asymmetric-configured systems at a scan rate of 50 mV s^{-1} in Fig. S2(a), the ReMIL-CN//AC device has a high area under the CV curve, indicating a higher charge storage capability and faster redox kinetics. This enhancement is attributed to the synergistic interaction between the transition metal selenide and the conductive $g\text{-C}_3\text{N}_4$ network, which facilitates efficient electron transfer and ion diffusion. Moreover, the nearly symmetrical redox peaks with a minimal potential shift at increasing scan rates reflect the excellent reversibility and rate capability of the composite. In contrast, the CV profiles of ReSe_2 //AC and ReSe_2 @MIL-53//AC cells show relatively smaller areas, implying slower reaction kinetics and a limited electroactive surface area. The GCD curves of ReSe_2 //AC and ReSe_2 @MIL-53//AC devices in Fig. S1(b and e) and GCD comparison plots of the three types of systems presented in Fig. S2(b) further corroborate the superior capacitive behavior of the hybrid electrode compared to its counterparts. All curves display nearly symmetric charge–discharge characteristics, indicating excellent reversibility and efficient charge-transfer kinetics. The discharge time decreases progressively with increasing current density, which can be

attributed to limited ion diffusion at higher rates. The smooth potential–time curves further confirm good electrical conductivity and stable capacitive behavior. The comparison results demonstrate that the ReMIL-CN//AC device possesses high rate capability and reliable electrochemical performance suitable for energy storage applications,⁶¹ with an extended discharge time, highlighting its enhanced energy storage capability. The excellent linearity and minimal IR drop observed in the GCD curves signify low internal resistance and high electrical conductivity. The improved electrochemical reversibility can be ascribed to the hierarchical architecture of the hybrid composite, where C_3N_4 sheets act as conductive bridges, preventing particle agglomeration and promoting rapid charge propagation throughout the electrode matrix. The specific capacity values calculated from the CV and GCD characteristic plots of ReSe_2 //AC and ReSe_2 @MIL-53//AC systems are presented in Fig. S1(c and f). The highest values of the specific capacities are 167.06 C g^{-1} for the ReSe_2 //AC system and 205.8 C g^{-1} for ReSe_2 @MIL-53//AC system. The ReSe_2 @ $g\text{-C}_3\text{N}_4$ //AC electrode-based system outperforms, which delivers high capacity, with a gradual less decrease as the scan rate increases in comparison to ReSe_2 //AC system. These measured values of the specific capacity are also lower than that of the hybrid electrode ReMIL-CN-based device (423 C g^{-1}). These findings confirm that the real ReMIL-CN//AC device outperforms the other developed composite

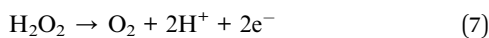
electrodes in terms of the specific capacity and electrochemical reversibility. The results collectively validate the effectiveness of our optimized ReMIL-CN composite-based system.

Based on the tunable efficiency and fast and high energy response of the present ReMIL-CN compound-based electrode, we have further evaluated its practical applicability in a hybrid device for the sensitive and selective detection of H₂O₂. It is a critical analyte in biomedical, food safety, and environmental monitoring applications, and its accurate and rapid detection is crucial, especially at trace levels.

3.3 Ternary ReMIL-CN composite-based electrode-integrated H₂O₂ detection system

For the H₂O₂ detection system configuration, the ReMIL-CN active electrode was tested based on its superior performance compared to other fabricated electrodes. The g-C₃N₄ provided a conducive platform for efficient electron transfer, the Fe-MOF introduced abundant active sites and porous channels to facilitate H₂O₂ diffusion and interaction, and ReSe₂ delivered high electrocatalytic activity for H₂O₂ oxidation.

Upon exposure to H₂O₂, the following electrocatalytic oxidation half-reaction occurred on the active electrode's surface:



The generated electrons were transmitted efficiently to the electrode, resulting in a measurable current that was proportional to the concentration of H₂O₂. This current was recorded using a potentiostat, forming the basis of the quantitative detection of H₂O₂. We aimed to assess the stability, reproducibility, and selectivity of the ReMIL-CN hybrid composite-based electrode under simulated real-world conditions. For this, the hybrid device was exposed to varying concentrations of H₂O₂ in complex sample matrices. Differential pulse voltammetry (DPV) and chronoamperometry (CA) measurements were employed for electrochemical H₂O₂ sensing. From these results, the response time, LOD, linear range (with *R*² value), and sensitivity *σ* of the sensor were quantified.

For electrochemical DPV analysis, a 0.1 M PBS electrolytic solution (with a pH value of 7.4) was used. The input scan rate potential was adjusted to 50 mV s⁻¹, corresponding to the effective output results, and the potential window was set from -0.2 V to +0.8 V vs. RHE (Hg/HgO). DPV curves were recorded at varying H₂O₂ concentrations ranging from 0 to 1000 μM (1 mM) with step values of 50, 100, 250, 500, 750, and 1000 μM. In the DPV plots presented in Fig. 12(a), clear anodic peaks were observed at 0.6 V vs. Hg/HgO, with peak currents increasing linearly as the H₂O₂ concentration increased. No noticeable redox peak was observed in the absence of H₂O₂ in the control test, as shown by the baseline current for 0 μM H₂O₂, clarifying the sensitivity and concentration-dependent response of the present modified electrode. Calibration curves obtained from the DPV plots shown in Fig. 12(b) demonstrated the linear dependence of peak current (*I*_p) on the H₂O₂ concentration values, highlighting the suitability of the presented ReMIL-CN nanocomposite electrode for the H₂O₂ detection system. The

sensitivity (*S*) of the device was measured from the slope of this calibration curve. The limit of detection (LOD) and limit of quantification (LOQ) were calculated using standard analytical equations, based on the standard deviation of the response and the slope of the calibration plot, as follows:

$$\text{LOD} = \frac{3\sigma}{S} \quad (8)$$

$$\text{LOQ} = \frac{10\sigma}{S} \quad (9)$$

Here, *S* is the sensitivity and *σ* corresponds to the standard deviation of the blank signal.

The DPV curve corresponding to 0 μM served as a baseline measurement in the absence of the H₂O₂ analyte. The ReMIL-CN nanocomposite's modified electrode demonstrated an immediate current response time of 3 s for each H₂O₂ addition, corresponding to 95% of the final current values. This minor response time demonstrated the good reproducibility of the electrode in the real-time H₂O₂ detection system. The linear regression value of *R*² = 0.989 presented a high linear response of the electrode for H₂O₂ sensing, with a linear operating range (LOR) from 50 μM to 1000 μM. For concentrations above 1000 μM, the signal deviated from linearity, which may be due to sensor saturation and the diffusion limit of the electrode. The key performance metrics measured for the presented electrode were a sensitivity (*S*) of 0.185 μA μM⁻¹ cm⁻², an LOQ of 0.333 μM, and a LOD of 0.1 μM.

For the CA analysis, the applied potential was set at 0.6 V. The H₂O₂ concentration ranged from 0 to 1000 μM with a step-wise rise of 50, 100, 250, 500, and 750 μM every 50 seconds. This increase in the concentration values was followed by continuous stirring to maintain its uniform diffusion. The obtained current vs. time curve demonstrated consistent current plateaus after each addition of H₂O₂, indicating the modified electrode's rapid and highly stable response. The remarkable sensing capability of the electrode resulted from the synergistic interactions within the composite material. ReSe₂ provided high electrocatalytic activity for H₂O₂ decomposition, the Fe-MOF offered abundant active sites and porous channels for H₂O₂ diffusion, and g-C₃N₄ ensured efficient electron transport and structural integrity.⁶² This combination enhanced sensitivity, fast response, and long-term stability, making the sensor suitable for real-time monitoring applications.

To elucidate the role of each component of the ReMIL-CN electrode in the H₂O₂ detection process, comparative electrochemical analyses were performed using ReSe₂, ReSe₂@MIL-53(Fe), and ReSe₂@MIL-53(Fe)@g-C₃N₄ electrodes. The DPV analysis plots of separate electrodes and comparative plots are shown in Fig. S3(a and c). The DPV peak potential and current intensity corresponded to the electrochemical reduction of H₂O₂ and varied notably depending on the structural composition, conductivity, and catalytic site distribution of each electrode material. Fig. S3(a) shows that the pristine ReSe₂ electrode exhibited a modest reduction current with a relatively positive peak potential, reflecting its moderate intrinsic conductivity and limited active site density. The layered



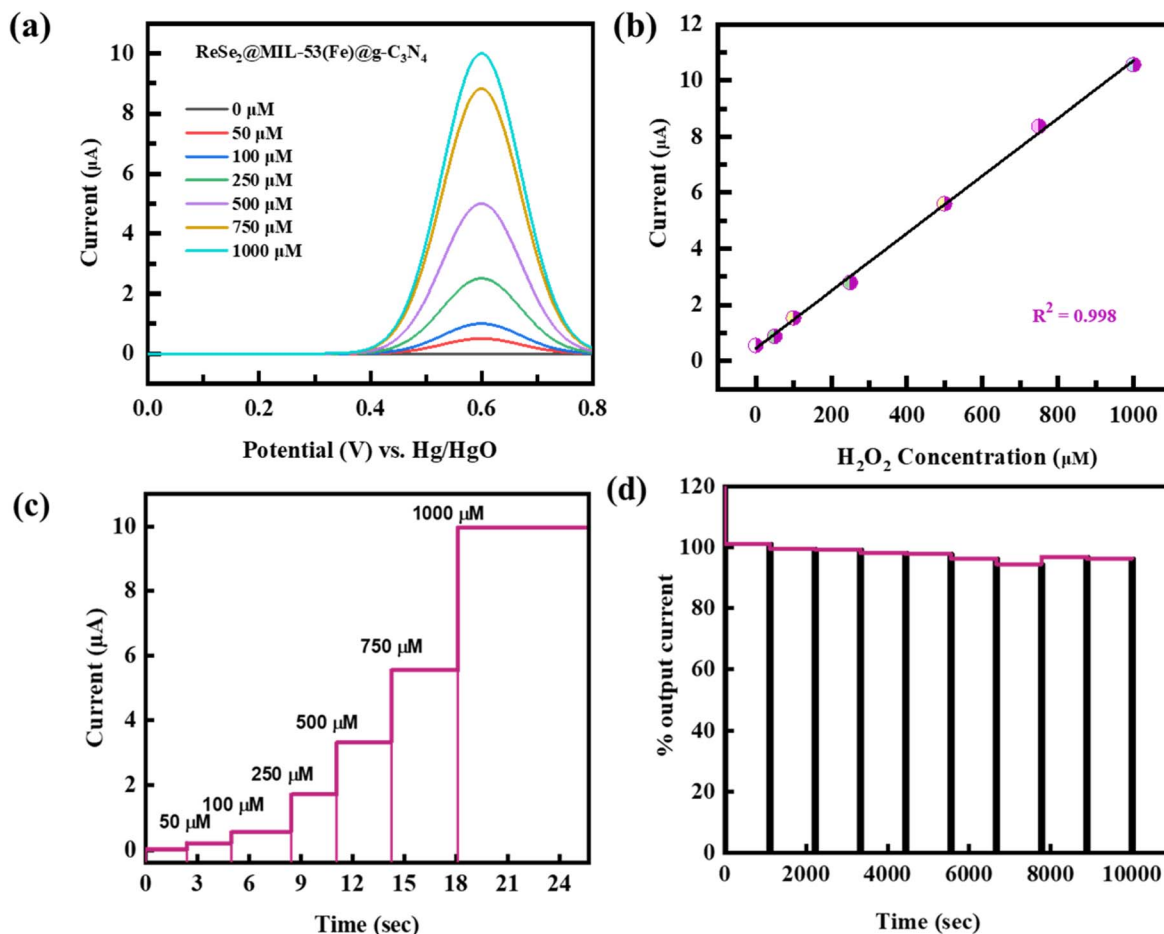


Fig. 12 (a) DPV curves at varying H₂O₂ concentrations, (b) calibration curve, (c) chronoamperometric (CA) evaluation of the response time, and (d) stability analysis for H₂O₂ sensing.

architecture and favorable Se-edge sites of ReSe₂ assisted charge transfer kinetics. The catalytic performance of the ReSe₂ electrode was somewhat hindered by the restricted number of reactive edges and weak adsorption affinity toward H₂O₂ molecules. The ReSe₂@MIL-53(Fe) electrode resulted in a noticeable enhancement in the catalytic current and a reduction in the onset potential. This is due to the Fe-based redox centers that promoted efficient electron mediation. The ReSe₂@MIL-53(Fe) electrode presented a distinctly enhanced DPV current response and a noticeable shift of the peak potential toward a more negative value. This enhancement facilitated the catalytic decomposition of H₂O₂ and served as an efficient electron mediator.⁶³ The porous MOF framework also increased the electrochemically active surface area and improved diffusion pathways for reactant molecules. The synergistic effect of the Fe nodes and the ReSe₂ conductive network promoted faster charge transfer and efficient utilization of active sites, as evidenced by the higher current density and reduced overpotential. The DPV plot comparison of ReSe₂, ReSe₂@MIL-53(Fe), and ReMIL-CN showed that the composite demonstrated the most pronounced catalytic behavior, characterized by the highest reduction peak current and the lowest overpotential. This improvement was primarily attributed to the

introduction of the g-C₃N₄ nanosheets, which provided abundant nitrogen functionalities that acted as electron-donating centers and facilitated H₂O₂ adsorption through N-H and lone-pair interactions.⁶⁴ Overall, the g-C₃N₄ supplied nitrogen-rich anchoring sites and improved surface adsorption and electronic coupling. The calibration plots of three types of electrodes demonstrated the linear rise in output currents with increasing H₂O₂ concentration. The excellent linearity confirmed the reliable sensitivity and reproducibility of the hybrid electrode. The DPV comparison plots in Fig. S4 clearly demonstrate that the ReMIL-CN plays an amplified overall catalytic performance through interfacial charge coupling and structural stabilization, rendering the composite an excellent candidate for high-performance H₂O₂ sensing applications.

To substantiate the performance and reliability of the fabricated electrode and its applied studies, a comprehensive comparison was performed, shown in Fig. 13 in which we have benchmarked our results against a wide range of previously reported systems for both BHSC and H₂O₂ detection applications, establishing a clear context for evaluating the superior efficiency, sensitivity, and practicality of the present work.

Fig. 13(a) illustrates the comparative performance of key metrics, including retention stability, energy density (E_d), and



making it usable in practical applications beyond the controlled laboratory settings, to evaluate matrix effects on sensor response in biological fluids and food extracts.

The presented comprehensive investigations highlighted the potential of the presented hybrid composite as a versatile multifunctional electrode material for dual-purpose electrochemical energy storage and hydrogen peroxide detection systems. Its dual functioning capability suggests a unique multifunctional advantage, where the electrode not only contributes to advanced energy storage but also enables real-time biochemical sensing. This feature positions it as a strong candidate for next-generation smart, self-powered bioelectrochemical devices and integrated health-monitoring supercapacitors, where energy storage and molecular detection functionalities coexist. It also paves the way for portable and wearable diagnostic technologies, food quality control, and environmental analysis. This synergistic approach not only enhances the practical utility of the electrode but also aligns with current trends in miniaturized, multifunctional, and wearable electronic devices.

Conflicts of interest

There are no conflicts of interest to declare.

Data availability

The data are available on request.

Supplementary information: Fig. S1 and S2: device level performance and comparison of ReSe₂//AC and ReSe₂@MIL-53//AC; Fig. S3 and S4: DPV curves, comparison and calibration plots of ReSe₂ and ReSe₂@MIL-53 composite. See DOI: <https://doi.org/10.1039/d5ra06541k>.

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