


Cite this: *RSC Adv.*, 2025, 15, 35930

Electronic and optical properties of halogen-substituted $\text{LaBi}_2\text{Cl}_{1-y}\text{X}_y\text{O}_4$: a promising candidate for energy-efficient devices

Dipendra Prasad Kalauni,^a Kedar Nath Jaiswal,^a Sarita Lawaju,^a Krishna Bahadur Rai,^a Ram Jeewan Yadav,^b Akkal Dev Mishra^c and Madhav Prasad Ghimire^{d,*ad}

Layered bismuth oxyhalides exhibit significant promise for photocatalytic and optoelectronic applications owing to their diverse functional characteristics. Here, we report the structural stability and electronic and optical properties of $\text{LaBi}_2\text{ClO}_4$ and its halogen-substituted derivatives $\text{LaBi}_2\text{Cl}_{1-y}\text{X}_y\text{O}_4$ ($\text{X} = \text{Br}, \text{I}$; $y = 0.25, 0.5, 0.75$) by means of a density functional theory approach. Both pristine and halogen-substituted compounds are thermodynamically stable, as evidenced by their negative formation energies, and the pristine phase satisfies the mechanical stability criteria. $\text{LaBi}_2\text{ClO}_4$ is identified as an indirect band gap semiconductor based on electronic structure calculations, with gap values between 1.18 eV (GGA) and 1.98 eV (mBJ). The dominant contributions near the valence and conduction band edges originate from O-2p and Bi-6p electronic states, respectively. Halogen substitution systematically narrows the band gap, with iodine doping inducing the most pronounced reduction, attributed to enhanced lattice distortions and modified orbital hybridization. Optical properties, computed via complex dielectric functions, exhibit pronounced anisotropy and tunability; doping shifts absorption edges toward lower photon energies and increases dielectric constants and refractive indices. These tunable optical responses lead to enhanced optical conductivity and reflectivity across the visible to ultraviolet spectrum, underscoring halogen-doped $\text{LaBi}_2\text{ClO}_4$ as a versatile platform for next-generation optoelectronic devices and photocatalysts.

Received 16th August 2025
Accepted 22nd September 2025

DOI: 10.1039/d5ra06049d

rsc.li/rsc-advances

1 Introduction

The development of semiconducting materials with tunable optical and electronic properties is critical for advancing next-generation technologies in renewable energy conversion,¹ environmental remediation,² and optoelectronic applications.^{3,4} In particular, materials capable of efficiently absorbing visible light, generating charge carriers, and facilitating redox reactions are central to addressing challenges in photocatalysis,^{5,6} solar energy harvesting,⁷ and pollutant degradation.⁶ Layered compounds have garnered significant attention in this context due to their intrinsic anisotropy, internal electric fields, and structural flexibility, which promote effective charge separation and enhance light–matter interactions.⁸ Among these, layered oxyhalides stand out as promising candidates, combining

oxide–halide frameworks with visible-light responsiveness and offering rich structural and chemical tunability.⁹ Bismuthoxyhalides (BiOX , $\text{X} = \text{Cl}, \text{Br}, \text{I}$) are frontier materials with layered structures, visible light responsiveness, and high charge-separation efficiency.^{10,11} Recent findings reveal outstanding performances in degrading antibiotics (e.g., BiOCl) and herbicides (e.g. BiOBr), but electron–hole recombination is still a problem.^{12,13} Semiconductor photocatalysis began in 1972 with the demonstration of water splitting using UV-irradiated TiO_2 (ref. 14). Since then, the field has grown rapidly, with major advances in efficient photocatalysts. Various semiconductors, including TiO_2 (ref. 15), ZnO ,¹⁶ ZrO_2 (ref. 17), SnO_2 (ref. 18) and Fe_2O_3 (ref. 19), have been used as photocatalysts. Among these semiconductors, titanium dioxide (TiO_2) has emerged as a prototypical n-type material, widely studied for its photocatalytic performance and electronic applications.^{14,20} Metal oxides, including TiO_2 , are vital for environmental cleanup as they create electron–hole pairs when exposed to adequate light energy.²¹ However, while TiO_2 is widely used in photocatalysis, it only works under UV light, which limits its effectiveness. In contrast, bismuth oxyhalides have an adjustable bandgap ranging from 3.3 eV to 1.8 eV for making them far more efficient at harnessing visible light for photocatalytic

^aCentral Department of Physics, Tribhuvan University, Kirtipur 44613, Kathmandu, Nepal. E-mail: madhav.ghimire@cdp.tu.edu.np

^bDepartment of Chemistry, Prithvi Narayan Campus, Tribhuvan University, Pokhara, Nepal

^cCentral Department of Chemistry, Tribhuvan University, Kirtipur 44613, Kathmandu, Nepal

^dLeibniz IFW Dresden, Helmholtzstr. 20, 01069 Dresden, Germany


reactions.^{22,23} BiOX (X = F, Cl, Br, I) has a layered structure with alternating $[\text{Bi}_2\text{O}_2]^{2+}$ positive layers and negatively charged halogen layers, generating an internal electric field. This promotes efficient electron-hole separation, enhancing photocatalytic performance in light-driven reactions.^{24,25} The band gap of bismuth oxyhalides can be modulated through multiple strategies, such as doping-induced band engineering, noble metal deposition, co-doping, and forming heterojunctions with narrow bandgap semiconductors, as widely documented in studies.²⁶ $\text{Bi}_2\text{REO}_4\text{X}$ (RE = Y, La–Lu; X = Cl, Br, I) forms a layered crystal structure, representing a combination of oxide and halide-based structural units with similarity in structure with BiOX. Fluorite-like $[\text{RE}_2\text{O}_2]$ layers alternate with single/double halide sheets in these intergrowths.^{27–29} Visible-light-responsive $\text{BiOCl}_x\text{Br}_{1-x}$, $\text{BiOCl}_x\text{I}_{1-x}$, and $\text{BiOBr}_x\text{I}_{1-x}$ with tunable E_g , VB, CB, and surface area (*via* x variation) show strong oxidation ability, making them effective photocatalysts for organic pollutant degradation.^{30,31} Previous studies have shown that the band gap of $\text{Bi}_2\text{REO}_4\text{X}$ (RE = Y, La–Lu) compounds can be tuned *via* doping, for instance in $\text{Bi}_2\text{Y}_{1-x}\text{M}_x\text{O}_4\text{Cl}$ (M = Fe, Ca, Zr).^{32,33} However, the structural and optoelectronic properties of the layered bismuth oxyhalide $\text{LaBi}_2\text{ClO}_4$, as well as the effects of halogen substitution at the Cl site with Br and I remain unexplored. In this work, we use density functional theory (DFT) to investigate the structural, mechanical, electronic and optoelectronic properties of $\text{LaBi}_2\text{ClO}_4$, and study the composition-dependent optoelectronic changes induced by Br and I substitution at varying concentrations. These insights extend halogen-tuning strategies to La-based oxyhalides and highlight their potential for optoelectronic applications.

2 Computational details

We carried out first-principles density functional theory (DFT) calculations to explore the structural, electronic, and optical properties of $\text{LaBi}_2\text{ClO}_4$ and its halogen-substituted variants $\text{LaBi}_2\text{Cl}_{1-y}\text{Br}_y\text{O}_4$ and $\text{LaBi}_2\text{Cl}_{1-y}\text{I}_y\text{O}_4$ at doping levels of $y = 0.25$, 0.5 , and 0.75 . Two computational tools were employed: the Full-Potential Local-Orbital (FPLO) code (v22.00-62)³⁴ and the Quantum ESPRESSO (QE v7.2) package.³⁵ In FPLO, the energy and charge convergence thresholds were chosen as 10^{-8} Hartree and $10^{-6}|e|$, respectively. Structural optimization involved k -point convergence testing followed by lattice parameter optimization. A $12 \times 12 \times 10$ k -point mesh was used for electronic structure calculations, while a denser $25 \times 25 \times 25$ k -mesh was employed for optical properties. The generalized gradient approximation (GGA) based on the PBE scheme³⁶ served as the exchange-correlation functional, and the modified Becke–Johnson (mBJ) potential³⁷ was employed to enhance band gap predictions. DFT provides reliable insights into electronic structures and related material properties, as demonstrated in previous studies.^{38,39} To simulate Br and I doping at the Cl site, supercells were constructed. The doped supercells retained the tetragonal crystal symmetry and the original space group. For $y = 0.25$ and $y = 0.75$, $2 \times 2 \times 1$ supercells containing four Cl atoms were used, enabling one- and three-site substitutions,

respectively. For $y = 0.5$, a more computationally efficient $1 \times 1 \times 2$ supercell with two Cl atoms was used, allowing 50% substitution while maintaining structural consistency. In all cases, multiple possible substitutional configurations were considered, and the ground-state energies were evaluated to identify the energetically most favorable doping arrangement, which was then used for further analysis. The mechanical properties were assessed using the quasi-harmonic approximation (QHA),⁴⁰ as incorporated in the thermo_pw code within the Quantum ESPRESSO package. The simulations employed a plane-wave basis set, applying cutoff values of 800 Ry for charge density and 80 Ry for kinetic energy. A stringent total energy convergence threshold of 10^{-8} Ry was applied to ensure computational accuracy.

3 Results and discussion

3.1 Structural stability

Bismuth lanthanum oxychloride, $\text{LaBi}_2\text{ClO}_4$, forms a tetragonal crystal structure belonging to the space group $P4/mmm$ (no. 123). A primitive unit cell, visualized using XFPLO as implemented in the FPLO package, is shown in Fig. 1(a). The atomic Wyckoff sites within the cell include: La located at (0.0, 0.0, 0.0), Bi at (−0.5, −0.5, 0.278377), Cl at (0.0, 0.0, −0.5), and O at (0.0, −0.5, 0.161656). To assess the structural stability of $\text{LaBi}_2\text{ClO}_4$, we performed a full structural relaxation by minimizing the total energy relative to the lattice parameters a and c using DFT within the GGA framework. The total energy difference $E - E_0$ (eV) was evaluated with respect to the lattice constants a and c , as illustrated in Fig. 1(b) and (c), respectively. Here, E_0 represents the minimum total energy corresponding to the equilibrium configuration. The lattice parameters were optimized by fitting the calculated total energy data to Murnaghan's equation of state,⁴¹ a commonly used method for describing how the energy of crystalline solids changes with volume. The resulting equilibrium lattice constants for $\text{LaBi}_2\text{ClO}_4$ are: $a = b = 4.03$ Å and $c = 9.33$ Å. These values show good consistency with, though slightly larger than, those reported within the Open Quantum Materials Database (OQMD), which provides $a = b = 3.964$ Å and $c = 9.064$ Å.⁴² A comparative summary of the optimized lattice parameters obtained from our calculations, alongside the values reported in the OQMD database and those derived from the $2 \times 2 \times 1$ and $1 \times 1 \times 2$ supercells, are presented in Table 1.

To evaluate the thermodynamic stability, we calculated the formation energy (E_f) of the parent compound $\text{LaBi}_2\text{ClO}_4$, obtaining a value of −1.82 eV per atom. The negative formation energy indicates that the compound is thermodynamically stable relative to its elemental constituents. This result compares well with the value of −2.26 eV per atom reported in the OQMD.⁴² The slight discrepancy may arise from differences in computational parameters or exchange–correlation functionals. Upon halogen substitution at the Cl site, the formation energies of the Br-doped compounds $\text{LaBi}_2\text{Cl}_{1-y}\text{Br}_y\text{O}_4$ are found to be −1.81, −1.80, and −1.79 eV per atom for doping levels y of 0.25, 0.50, and 0.75, respectively. In comparison, I-doped counterparts $\text{LaBi}_2\text{Cl}_{1-y}\text{I}_y\text{O}_4$ exhibit slightly less negative

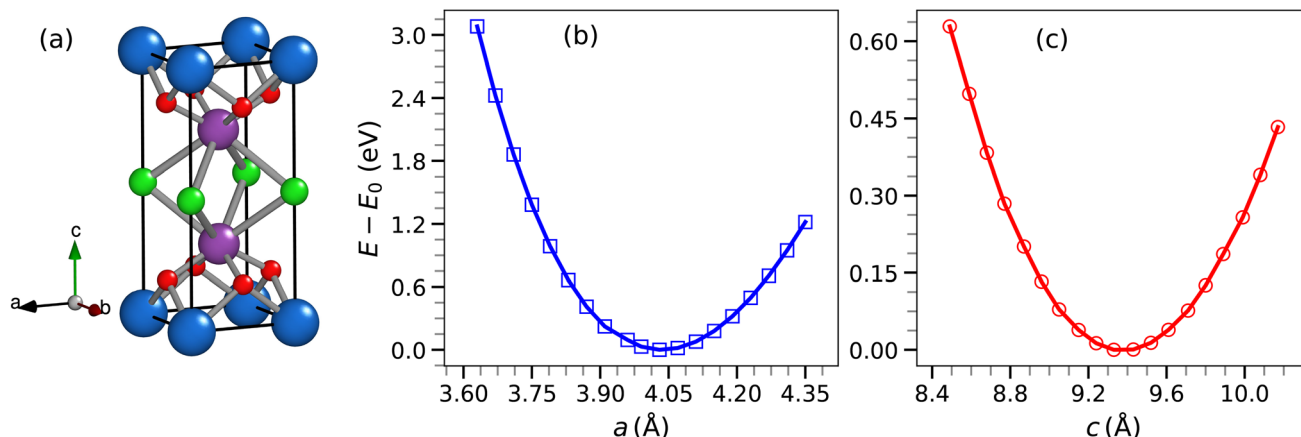


Fig. 1 (a) Tetragonal atomic arrangement of $\text{LaBi}_2\text{ClO}_4$, where blue, violet, green, and red spheres represent La, Bi, Cl, and O atoms, respectively. (b) Relative total energy with respect to lattice parameter a (in Å). (c) Relative total energy, $(E - E_0)$, with respect to the lattice parameter c (in Å). The curves in panels (b) and (c) are fitted using Murnaghan's equation of state.

formation energies of -1.79 , -1.76 , and -1.73 eV per atom for the same doping levels. These results indicate that all compounds remain thermodynamically stable upon doping, with Br substitution yielding slightly more favorable formation energies than I substitution.

3.2 Mechanical properties

The elastic properties of $\text{LaBi}_2\text{ClO}_4$ were systematically investigated to evaluate its mechanical behavior and stability. The independent elastic constants C_{ij} , derived from the generalized Hooke's law,⁴³ serve as fundamental parameters linking atomic bonding characteristics to the macroscopic mechanical response of the material. For the tetragonal phase (space group $P4/mmm$) of $\text{LaBi}_2\text{ClO}_4$, six distinct elastic constants— C_{11} , C_{12} , C_{13} , C_{33} , C_{44} , and C_{66} —were calculated using the QHA method and are shown in Table 2.

Mechanical properties of a material provide critical insights into its response to external stresses and are key indicators of its structural stability. The mechanical stability of such crystals is governed according to the Born criteria. In the case of a tetragonal system classified as type I,⁴⁴ the structure is mechanically stable if the following conditions are satisfied: $C_{11} > |C_{12}|$, $C_{33} > 0$, $C_{44} > 0$, $C_{66} > 0$, and $2C_{13}^2 < C_{33}(C_{11} + C_{12})$. The computed elastic constants for $\text{LaBi}_2\text{ClO}_4$, listed in Table 2, satisfy all of these conditions, confirming that the compound is mechanically stable. The elastic response of $\text{LaBi}_2\text{ClO}_4$ was further characterized through the evaluation of macroscopic mechanical moduli derived from the computed elastic constants. The bulk modulus ($B = 62.32$ GPa) reflects the material's resistance

to isotropic compression, indicating moderate incompressibility under hydrostatic pressure. The shear modulus ($G = 26.72$ GPa), which governs the material's resistance to shape deformation under shear stress, is relatively low, consistent with the layered structure of the compound.⁴⁵ Young's modulus ($Y = 69.82$ GPa), which quantifies stiffness under uniaxial strain,^{46,47} also supports the conclusion that $\text{LaBi}_2\text{ClO}_4$ exhibits moderate elastic rigidity. The Pugh's ratio⁴⁸ ($B/G = 2.33$) significantly exceeds the empirical ductile–brittle threshold of 1.75, indicating pronounced ductility. This is further corroborated by its inverse ($G/B = 0.43$), which lies well below unity, reaffirming the ductile mechanical character. The calculated Poisson's ratio ($\nu = 0.31$) falls within the typical range for ionic solids, suggesting an intermediate bonding character with a dominant ionic contribution.⁴⁹ Additionally, a Poisson's ratio close to 0.5 implies low volume change under stress; thus, the observed value suggests that $\text{LaBi}_2\text{ClO}_4$ possesses relatively low compressibility and good mechanical stability. These findings—moderate stiffness, high ductility, and ionic bonding character—are consistent with previously reported mechanical behavior of structurally analogous layered bismuth oxyhalides, such as $\text{Bi}_2\text{LaO}_4\text{I}$.⁵⁰

3.3 Electronic properties

In order to determine the ground-state properties of $\text{LaBi}_2\text{ClO}_4$, both nonmagnetic and spin-polarized configurations were investigated. The total energies of the two configurations were found to be identical within the convergence threshold, and the spin-polarized calculations yielded zero magnetic moments,

Table 1 Optimized lattice constants $a = b$ and c (in Å) for $\text{LaBi}_2\text{ClO}_4$ obtained from our DFT calculations, compared with OQMD data⁴² and results from $2 \times 2 \times 1$ and $1 \times 1 \times 2$ supercell configurations

Compound	$a = b$	c	$a_{\text{ref}} = b_{\text{ref}}$	c_{ref}	$a_{2 \times 2 \times 1} = b_{2 \times 2 \times 1}$	$c_{2 \times 2 \times 1}$	$a_{1 \times 1 \times 2} = b_{1 \times 1 \times 2}$	$c_{1 \times 1 \times 2}$
$\text{LaBi}_2\text{ClO}_4$	4.03	9.33	3.964	9.064	8.06	9.20	4.03	18.66



Table 2 Calculated elastic constants (C_{ij} in GPa), Young's modulus (Y in GPa), shear modulus (G in GPa), bulk modulus (B in GPa), Pugh's ratio (B/G), its inverse (G/B), and Poisson's ratio (ν) for the layered bismuth oxyhalide $\text{LaBi}_2\text{ClO}_4$

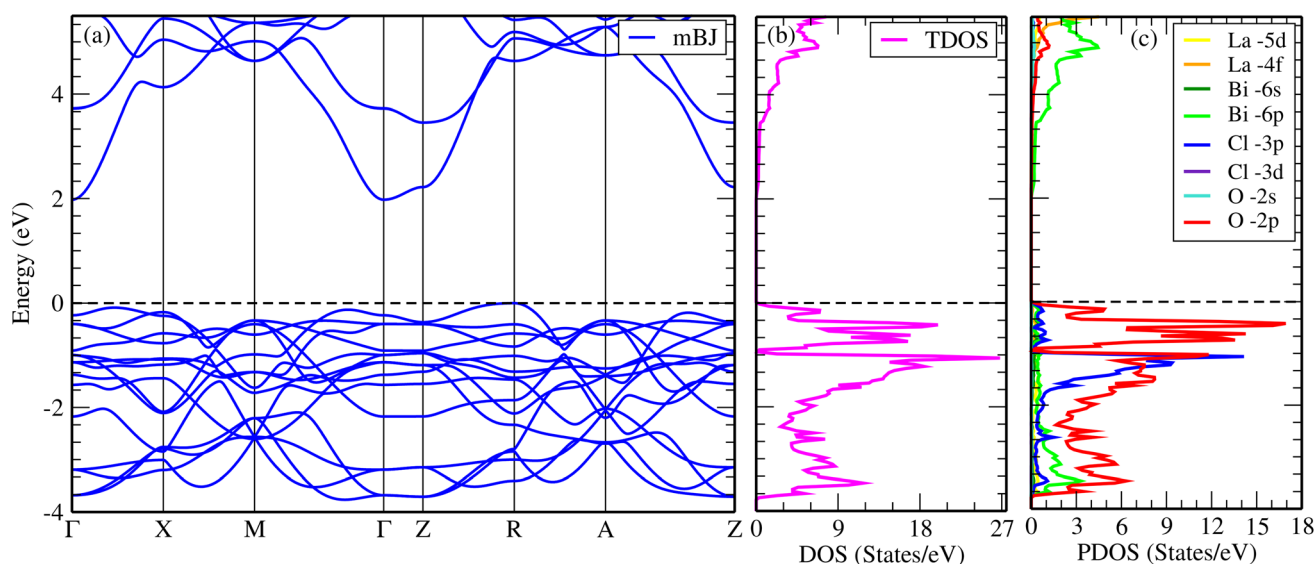
Material	C_{11}	C_{12}	C_{13}	C_{33}	C_{44}	C_{66}	Y	G	B	B/G	G/B	ν
$\text{LaBi}_2\text{ClO}_4$	146.64	64.04	25.41	78.85	10.41	63.79	69.82	26.72	62.32	2.33	0.43	0.31

indicating a nonmagnetic ground state. To investigate the electronic configuration of $\text{LaBi}_2\text{ClO}_4$, density functional theory calculations were carried out employing both GGA and mBJ exchange potentials. The overall features of the electronic band structure remain similar in both approaches, except for a significant increase in the band gap observed with the mBJ functional. We focus on mBJ calculations for the electronic band structure because the GGA approach underestimates band gaps,^{51,52} whereas mBJ closely reproduces experimental trends, as shown by the agreement between measured and calculated band gaps in this class of oxyhalides,³² demonstrating its reliability for accurately predicting their electronic properties. For comparison, the GGA results are provided in the SI, while this section emphasizes the findings obtained using mBJ, with occasional references to GGA where relevant.

Fig. 2(a) illustrates the electronic band structure for $\text{LaBi}_2\text{ClO}_4$ obtained using the mBJ approach. The valence band maximum (VBM) occurs at the high-symmetry point R , while the conduction band minimum (CBM) appears at the Γ point, confirming that $\text{LaBi}_2\text{ClO}_4$ exhibits an indirect band gap.⁵³ Such an indirect transition is advantageous for reducing the recombination probability of photoinduced carriers, as the photoexcited electrons must traverse a finite distance in k -space before recombination occurs.^{54,55} Using the GGA method, the indirect band gap is calculated as 1.18 eV (see SI: Fig. S1(a)), whereas the mBJ approach yields a significantly larger value of 1.98 eV, consistent with the well-known band gap underestimation

tendency of GGA.⁵¹ For $\text{LaBi}_2\text{ClO}_4$ to be effective in photoelectronic applications, both the size and the nature of its band gap are of great importance. The band gap size influences how the material interacts with sunlight across the spectrum, while the gap type governs charge carrier behavior. Materials with a direct band gap typically absorb photons more efficiently, whereas those with an indirect band gap tend to promote better separation of photoexcited electrons and holes, thereby reducing their likelihood of recombination.⁵⁶ This behavior aligns well with previously reported findings on structurally related layered bismuth oxyhalides, including $\text{Bi}_2\text{LaO}_4\text{I}^{50}$ and $\text{Bi}_2\text{YO}_4\text{Cl}$,³² which shows comparable electronic and crystallographic features.

Fig. 2(b) and (c) depict the total and partial density of states (DOS and PDOS) of $\text{LaBi}_2\text{ClO}_4$, respectively. As observed in the total DOS, a significant band gap of around 1.18 eV appears using the GGA approach, while a larger gap of 1.98 eV is found with the mBJ functional, consistent with the results of the electronic band structure. The orbital-resolved PDOS shown in Fig. 2(c) (see SI: Fig. S1(c)) reveals that the valence band states close to the Fermi level (E_F) are primarily derived from O-2p, Cl-3p, and Bi-6s orbitals, whereas the conduction band region largely consists of Bi-6p, La-4f states, and La-5d contributions. La-5d, La-4f, and Bi-6s states contribute minimally to the valence band, while Cl-3d, Cl-3p, and O-2s orbitals exhibit negligible presence in the conduction band region. There is a marked hybridization between the Cl-3p and O-2p orbitals in the valence band, revealing substantial p-p interaction in this

**Fig. 2** Calculated electronic properties of $\text{LaBi}_2\text{ClO}_4$ using the mBJ potential: (a) band structure, (b) total density of states (DOS), and (c) projected density of states (PDOS). Dashed horizontal lines in all plots denote the Fermi level (E_F).

area. In contrast, a modest degree of hybridization is present between Cl-3p and Bi-6p orbitals within the conduction band. Furthermore, Bi-6p and La-5d orbitals contribute predominantly to the conduction band states, particularly at higher energies above the Fermi level, although their mutual hybridization remains moderate. These orbital characteristics give rise to a wide band gap and a favorable electronic structure, rendering $\text{LaBi}_2\text{ClO}_4$ a promising material for optoelectronic applications such as photocatalysis, photodetection, and light-emitting diodes (LEDs).⁴

Fig. 3(a–c) displays the electronic band structures of $\text{LaBi}_2\text{Cl}_{1-y}\text{Br}_y\text{O}_4$, while Fig. 3(d–f) shows those of $\text{LaBi}_2\text{Cl}_{1-y}\text{I}_y\text{O}_4$ for doping levels $y = 0.25, 0.50$, and 0.75 , where Cl atoms are partially substituted by Br and I, respectively. To model 25% doping, one Cl atom out of four was substituted in a $2 \times 2 \times 1$ supercell. For 50% doping, one Cl atom out of two was replaced in a $1 \times 1 \times 2$ supercell. In the case of 75% doping, three Cl atoms out of four were substituted in a $2 \times 2 \times 1$ supercell. In all doped configurations, the CBM and VBM appear at different high-symmetry points, confirming that the systems remain indirect band gap semiconductors, consistent with pristine $\text{LaBi}_2\text{ClO}_4$ (band gap: 1.18 eV, GGA; 1.98 eV, mBJ; VBM at R, CBM at Γ). At $y = 0.25$, replacing Cl with Br results in an indirect band gap of 1.13 eV (GGA) and 1.92 eV (mBJ). A similar substitution with iodine yields slightly lower values of 1.07 eV (GGA) and 1.83 eV (mBJ). These values show a decrease compared to pristine $\text{LaBi}_2\text{ClO}_4$, indicating the pronounced influence of heavier halogen atoms on the material's electronic structure. The initial reduction in band gap arises from the larger atomic radii of Br and I,^{57–59} which expand the lattice and perturb the crystal field. With increased substitution ($y = 0.50$), this trend continues: Br doping yields band gaps of 1.08 eV (GGA) and 1.92 eV (mBJ), while I doping further lowers the gap

to 1.02 eV (GGA) and 1.70 eV (mBJ). At $y = 0.75$, the Br-substituted compound shows values of 1.05 eV (GGA) and 1.85 eV (mBJ), whereas I doping reduces the band gap to 0.96 eV (GGA) and 1.64 eV (mBJ). These findings reveal a monotonic decrease in the indirect band gap with progressive halogen substitution. This effect is attributed to the increased ionic radii of Br and I, which induce local lattice distortions and modify the bonding environment. As a result, these structural changes affect orbital hybridization and shift the energy positions of the valence and conduction bands. Such tunability of the electronic properties underscores the potential of halogen-substituted layered oxyhalides for optoelectronic⁶⁰ and photocatalytic applications,^{9,32} where precise control over the band gap is essential. Fig. 4(a–f) shows the DOS and PDOS of $\text{LaBi}_2\text{Cl}_{1-y}\text{Br}_y\text{O}_4$ at substitution levels of $y = 0.25, 0.50$, and 0.75 . The total DOS plots (Fig. 4(a–c)) reveal a gradual narrowing of the band gap with increasing Br concentration, decreasing from approximately 1.13 to 1.05 eV based on the GGA results (see SI: Fig. S3(a–c)). The modified Becke–Johnson (mBJ) potential, in contrast, yields higher band gap values in the range of 1.92 to 1.81 eV, which align well with the electronic band structures illustrated in Fig. 3(a–c). The PDOS plots (Fig. 4(d–f)) reveal that the states in the vicinity of the Fermi energy (E_F) within the valence band region are predominantly derived from O-2p and Cl-3p orbitals, while Br-4p contributions become more pronounced as doping increases. The CBM is primarily composed of Bi-6p states. Lesser yet notable contributions in the conduction region come from La-5d, La-4f, and O-2s orbitals, with moderate involvement of Bi-6s and Bi-6p states in the valence region. Orbitals such as La-5d, La-4f, and Bi-6s (valence) and Cl-3d, Cl-3p, and O-2s (conduction) exhibit minimal involvement. Distinct hybridization among Cl-3p, O-2p, and Br-4p orbitals is evident in the valence band,

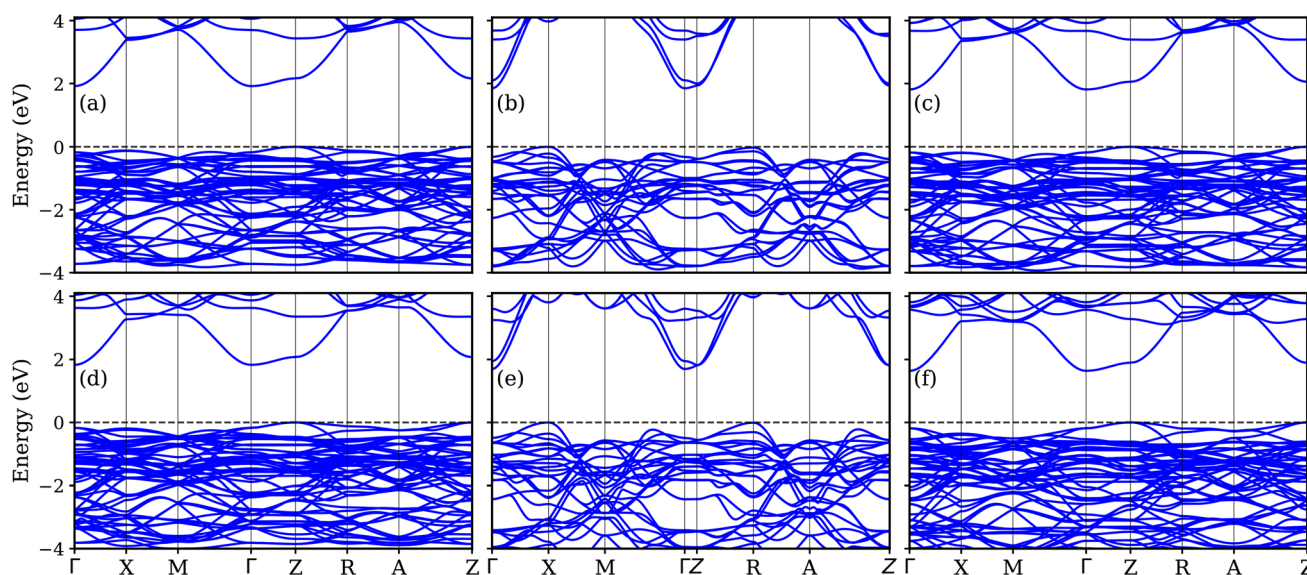


Fig. 3 Electronic band structures of $\text{LaBi}_2\text{ClO}_4$ with partial substitution of Cl by Br and I at various doping concentrations, calculated using the mBJ potential: (a) $\text{LaBi}_2\text{Cl}_{0.75}\text{Br}_{0.25}\text{O}_4$, (b) $\text{LaBi}_2\text{Cl}_{0.5}\text{Br}_{0.5}\text{O}_4$, (c) $\text{LaBi}_2\text{Cl}_{0.25}\text{Br}_{0.75}\text{O}_4$, (d) $\text{LaBi}_2\text{Cl}_{0.75}\text{I}_{0.25}\text{O}_4$, (e) $\text{LaBi}_2\text{Cl}_{0.5}\text{I}_{0.5}\text{O}_4$, and (f) $\text{LaBi}_2\text{Cl}_{0.25}\text{I}_{0.75}\text{O}_4$.



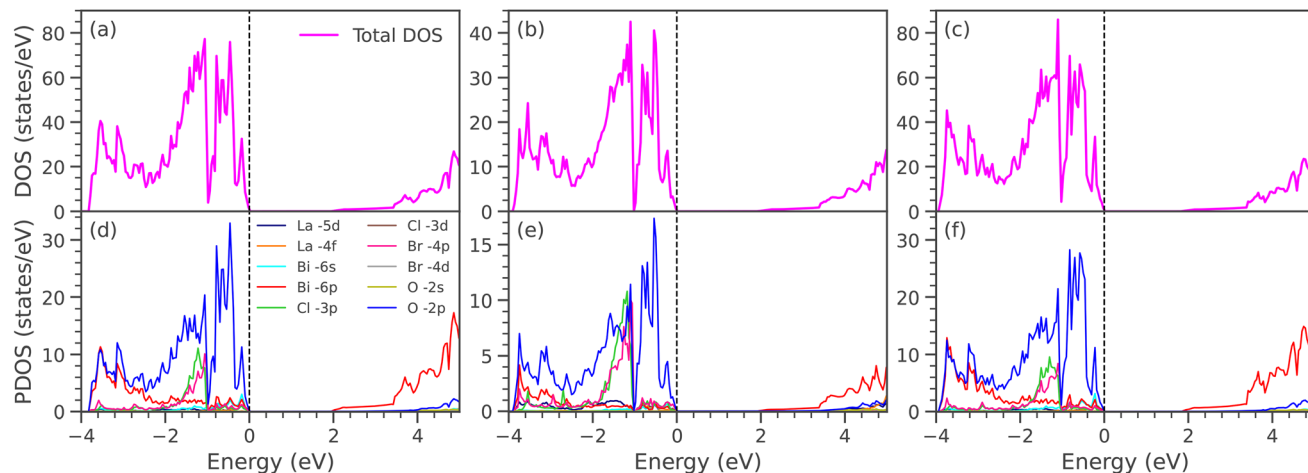


Fig. 4 Total and partial DOS of (a and d) LaBi₂Cl_{0.75}Br_{0.25}O₄, (b and e) LaBi₂Cl_{0.5}Br_{0.5}O₄, (c and f) LaBi₂Cl_{0.25}Br_{0.75}O₄, calculated using the mBJ potential.

alongside intermediate hybridization between Cl-3p and Bi-6p. Additionally, weak interactions are identified between Bi-6s and O-2p, and moderate hybridization occurs between O-2s and Bi-6p in the conduction band region. Fig. 5(a–f) displays the total and orbital-resolved DOS for LaBi₂Cl_{1–y}I_yO₄ at the same doping levels. As depicted in the total DOS (Fig. 5(a–c)) a progressive narrowing of the GGA band gap occurs from 1.07 to 0.96 eV with increasing I content (see SI: Fig. S3(d–f)). The mBJ-calculated gaps span a higher range, decreasing slightly from 1.83 to 1.64 eV, in line with the corresponding band structures (Fig. 3(d–f)). According to the PDOS (Fig. 5(d–f)), the valence states near E_F primarily originate from O-2p and Cl-3p orbitals, while I-5p orbitals exhibit growing intensity with increasing substitution. Bi-6p states remain the dominant contributors at the CBM throughout the doping range. Furthermore, La-5d, La-4f, and O-2s orbitals make modest contributions in the conduction band, while Bi-6s and Bi-6p are visible in the valence region. Orbitals such as La-5d, La-4f, and Bi-6s (valence), along with Cl-3d, Cl-3p, and O-2s (conduction), have

negligible impact. Prominent hybridization is noted among Cl-3p, O-2p, and I-5p orbitals in the valence region, complemented by moderate Cl-3p–Bi-6p interactions. Weak coupling between Bi-6s and O-2p, as well as moderate O-2s–Bi-6p hybridization in the conduction band, is also observed. Compared with co-doping and heterostructure-based strategies, halogen substitution offers a simpler and more reproducible approach for band-gap engineering. Co-doping, such as Al–Co in TiO₂ (ref. 61) or La–N in SrTiO₃,⁶² achieves comparable reductions in the electronic gap (3.11 eV → 1.97 eV and 2.20 eV → 1.02 eV) but requires multi-dopant incorporation, which can introduce defect states. Heterostructures, including ZnO/BSe⁶³ and HfS₂/MoTe₂,⁶⁴ enable precise band alignment but demand complex interface engineering. In contrast, single-site halogen doping is straightforward, though very high substitution levels may induce lattice strain or defects. This tunable strategy positions LaBi₂ClO₄ as a promising candidate for optoelectronic and photocatalytic applications.

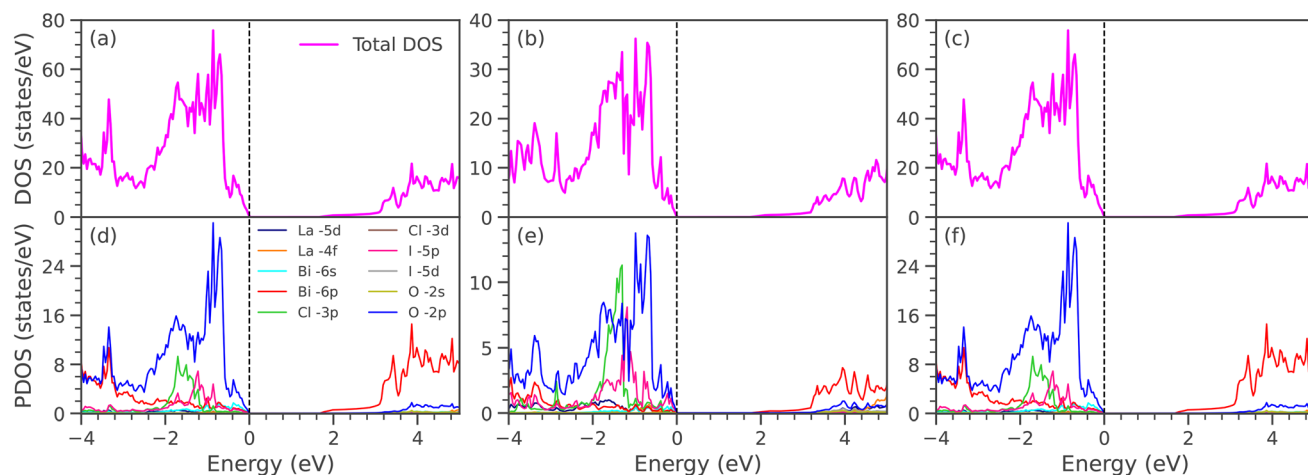


Fig. 5 Total and partial DOS of (a and d) LaBi₂Cl_{0.75}I_{0.25}O₄, (b and e) LaBi₂Cl_{0.5}I_{0.5}O₄, and (c and f) LaBi₂Cl_{0.25}I_{0.75}O₄ calculated using the mBJ potential.

3.4 Optical properties

Fig. 6(a–h) illustrates the optical properties of $\text{LaBi}_2\text{ClO}_4$ computed using the mBJ exchange potential. These include (a) the real and (b) imaginary parts of the dielectric function, (c) energy loss function, (d) optical conductivity, (e) refractive index, (f) extinction coefficient, (g) absorption coefficient, and (h) reflectivity. The xx and zz components refer to the polarization along the x - and z -axes, respectively. For comparison, the corresponding optical properties were also calculated using the GGA functional. The results exhibit the same overall behavior as mBJ, with only slight differences in magnitude (see SI for details).

The complex dielectric function is a fundamental quantity used to characterize the optical response of materials.^{65–67} It describes how a material interacts with incident electromagnetic radiation and governs various frequency-dependent optical properties. The real part of the dielectric function, $\epsilon_1(\omega)$, reflects the material's polarization response and dispersion behavior, while the imaginary part, $\epsilon_2(\omega)$, is directly related to the absorption of electromagnetic radiation and reveals information about interband electronic transitions. As such, the dielectric function serves as a central parameter through which other optical constants can be derived and analyzed. Fig. 6(a) presents the real part of the dielectric function, $\epsilon_1(\omega)$, in terms of photon energy ranging from 0 to 12 eV. The static dielectric constant is given by $\epsilon_1(0)$ at zero frequency, with calculated values of 3.56 and 3.30 along the crystallographic directions xx and zz , respectively. The static dielectric function indicates the degree of charge recombination in a material, where a higher value corresponds to a lower recombination rate, thereby playing a crucial role in enhancing the operational effectiveness of optoelectronic devices. When the value moves into the UV spectrum, it begins to decline slowly, causing the real part of the dielectric function to take on negative values within this range.⁶⁸ Between 6.59 and 10.1 eV, the real part of the dielectric function for the compound exhibits negative values. Within this energy range, the dielectric function repeatedly attains negative values due to multiple peaks in the

visible and UV regions. With increasing photon energy, $\epsilon_1(\omega)$ increases, reaching maxima of 7.92 at 5.07 eV (xx) and 7.54 at 6.19 eV (zz). Beyond these energies, $\epsilon_1(\omega)$ decreases and turns negative, a behavior commonly linked to interband transitions that align with significant peaks in the dielectric function's imaginary component.

Variation of the imaginary component of the dielectric function, $\epsilon_2(\omega)$, with respect to photon energy is depicted in Fig. 6(b). Peaks in $\epsilon_2(\omega)$ reflect transitions of electrons between the valence and conduction bands and indicate the energy levels at which photons are absorbed. These absorption characteristics reveal details about the compound's interaction with incoming radiation. The optical absorption begins at around 3.65 eV, indicating the optical band gap of $\text{LaBi}_2\text{ClO}_4$. Notable anisotropy is observed between the xx and zz polarization directions. Distinct absorption peaks appear at 5.17, 5.71, 6.33, 6.99, and 7.59 eV for the xx polarization, and at 5.93, 6.37, and 6.62 eV for the zz polarization. These features correspond to interband transitions, primarily from the O 2p valence band to the Bi 6p conduction band, which are consistent with the trends observed in $\epsilon_1(\omega)$. The energy dissipated by a fast electron passing through the material is characterized by the energy-loss function, $L(\omega)$, as presented in Fig. 6(c).⁶⁹ The peaks in $L(\omega)$ are typically associated with plasmon excitations. A pronounced peak is observed at approximately 10.7 eV along the xx -direction, indicating significant energy loss in this polarization. In contrast, the peak along the zz -direction appears at a lower energy of around 8.5 eV, highlighting the anisotropic plasmonic response of $\text{LaBi}_2\text{ClO}_4$. The optical conductivity, $\sigma(\omega)$, shown in Fig. 6(d), describes the process in which electrons are excited from energy states within the valence band into available states in the conduction band through photon absorption.⁶⁹ It follows a similar profile to that of $\epsilon_2(\omega)$, as they are related by the expression $\sigma(\omega) = \frac{\omega}{4\pi} \epsilon_2(\omega)$. Several sharp features are observed in the conductivity spectrum, which may arise from interband transitions and plasmonic excitations. The maximum values of $\sigma(\omega)$ along the xx and zz polarization directions are approximately 6.9 and 8.6 ($\times 10^3 \Omega^{-1} \text{ cm}^{-1}$), at photon energies of 6.31

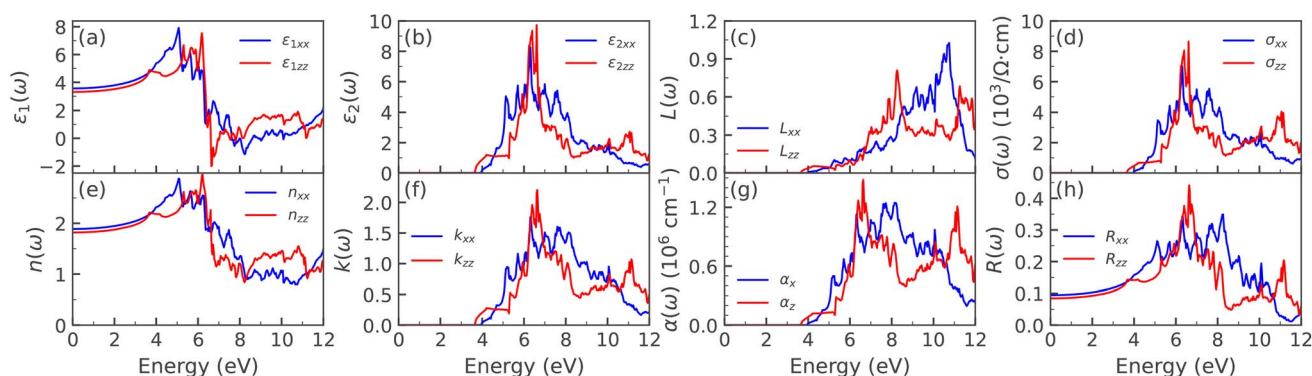


Fig. 6 Calculated optical properties of $\text{LaBi}_2\text{ClO}_4$ as functions of photon energy using the mBJ exchange potential. Panels show: (a) real $\epsilon_1(\omega)$ and (b) imaginary $\epsilon_2(\omega)$ parts of the dielectric function; (c) energy loss function $L(\omega)$; (d) optical conductivity $\sigma(\omega)$; (e) refractive index $n(\omega)$; (f) extinction coefficient $k(\omega)$; (g) absorption coefficient $\alpha(\omega)$; and (h) reflectivity $R(\omega)$. The subscripts xx and zz denote the polarization directions along the crystallographic x - and z -axes, respectively.



and 6.62 eV, respectively. Fig. 6(e) illustrates how the refractive index, $n(\omega)$, of $\text{LaBi}_2\text{ClO}_4$ varies with photon energy. The refractive index is a fundamental optical parameter that quantifies the reduction in the speed of light within a medium and governs the refraction of light at material interfaces.⁷⁰ Materials with a high refractive index—typically exceeding 1.8—are particularly desirable for optoelectronic and photovoltaic applications⁶⁶ due to their enhanced light–matter interaction capabilities. As evident from Fig. 6(e), $\text{LaBi}_2\text{ClO}_4$ exhibits a static refractive index of approximately 1.88 and 1.81 along the xx and zz directions, respectively, at zero photon energy. The refractive index attains peak values of 2.87 in the xx direction and 2.96 in the zz direction at around 5.13 eV and 6.19 eV photon energies. This relatively high refractive index implies strong photon confinement and an enhanced optical response, properties that are advantageous for applications in solar cells, photodetectors, and light-emitting diodes.⁴ The absorption and scattering behavior of electromagnetic radiation in the material, as a function of photon energy, is described by the extinction coefficient $k(\omega)$.⁷¹ Fig. 6(f) illustrates that $\text{LaBi}_2\text{ClO}_4$ exhibits a notable rise in $k(\omega)$ near 6.5 eV, reaching a maximum value of approximately 2.25, which indicates strong optical absorption in this energy range. The absorption coefficient quantifies the optical energy loss in a material as light passes through a specific length.⁷² Optical absorption is triggered when the photon frequency coincides with electronic transitions inside the material. The absorption coefficient, $\alpha(\omega)$, varies with photon energy, reflecting the material's ability to absorb light at specific frequencies. When photons possess sufficient energy, electrons are excited from the valence band to the conduction band, resulting in optical absorption. Fig. 6(g) presents the photon-energy-dependent absorption coefficient, $\alpha(\omega)$, of $\text{LaBi}_2\text{ClO}_4$. A distinct absorption edge appears near ~ 3.65 eV, marking the onset of interband transitions. Beyond this threshold, $\alpha(\omega)$ increases significantly, and several pronounced peaks are observed between 5.0 and 11.5 eV, indicating strong optical activity and photon absorption in this energy range. The reflectivity spectrum, $R(\omega)$, is presented in Fig. 6(h). The maximum reflectivity occurs near 6.33 and 6.65 eV, with peak values of 34% and 43% along the xx - and zz -directions, respectively. The relatively low reflectivity throughout the examined energy range indicates strong optical absorption in $\text{LaBi}_2\text{ClO}_4$, with minimal energy loss due to reflection, thereby suggesting its potential suitability for optoelectronic applications.⁵³ The influence of halide substitution on the optical properties of $\text{LaBi}_2\text{ClO}_4$ was systematically investigated for Br and I doping at concentrations of $y = 0.25, 0.50$, and 0.75 . Fig. 7(a)–(h) and 8(a)–(h) present the frequency-dependent optical parameters of $\text{LaBi}_2\text{Cl}_{1-y}\text{Br}_y\text{O}_4$ and $\text{LaBi}_2\text{Cl}_{1-y}\text{I}_y\text{O}_4$, respectively. Fig. 7(a) and 8(a) present the real part of the dielectric function, $\epsilon_1(\omega)$. The static dielectric constant, $\epsilon_1(0)$, increases with substitution in both Br- and I-doped systems. Compared to the parent compound (3.56/3.30 along the xx/zz directions), Br substitution leads to moderate increases, reaching 3.68/3.40 for $y = 0.25$, 3.79/3.50 for $y = 0.50$, and 3.91/3.60 for $y = 0.75$. For I substitution, the increase is more pronounced, with $\epsilon_1(0)$ values of 4.09/3.73 for $y = 0.25$, 4.71/4.15

for $y = 0.50$, and 5.24/4.53 for $y = 0.75$, indicating enhanced polarizability due to the heavier halogen. Additionally, the main peaks in $\epsilon_1(\omega)$ shift to higher energies with increasing y , in contrast to the peak positions at 5.07 eV (xx) and 6.19 eV (zz) in the undoped material. Fig. 7(b) and 8(b) show the imaginary part, $\epsilon_2(\omega)$. The optical absorption onset redshifts with doping. In the Br-doped case, it shifts from 3.42 eV ($y = 0.25$) to 2.75 eV ($y = 0.75$); for I doping, a more substantial shift occurs—from 1.93 eV ($y = 0.25$) to 1.47 eV ($y = 0.75$). These features indicate narrowing band gaps and are attributed to interband transitions involving O 2p to Bi 6p states. The shift is more prominent in I-doped compounds, demonstrating stronger modification of the optical band structure. The energy-loss function $L(\omega)$ (Fig. 7(c) and 8(c)) exhibits pronounced plasmon peaks. In the xx direction, the plasmon peak remains near 10.7 eV for all Br and I concentrations, comparable to the undoped system. Along the zz direction, Br substitution slightly increases the peak energy from 8.17 eV at $y = 0.25$ to 8.31 eV at $y = 0.50$, while it remains relatively unchanged at $y = 0.75$. In contrast, I doping leads to a downward trend in peak positions, decreasing from 8.17 eV at $x = 0.25$ to 7.43 eV at $y = 0.75$, indicating a softened plasmonic response and enhanced anisotropy in dielectric behavior. Fig. 7(d) and 8(d) present the optical conductivity, $\sigma(\omega)$, in units of $\Omega^{-1} \text{ cm}^{-1}$, for Br- and I-doped $\text{LaBi}_2\text{ClO}_4$. In Br-doped systems, peak values increase with doping, especially along the zz direction: for $y = 0.25$, $\sigma(\omega)$ reaches 5.7×10^3 (xx , at 6.22 eV) and 7.8×10^3 (zz , at 6.31 eV); at $y = 0.50$, 5.3×10^3 (xx) and 8.9×10^3 (zz); and at $y = 0.75$, 5.1×10^3 (xx) and 9.1×10^3 (zz). Iodine doping follows a non-monotonic trend: at $y = 0.25$, $\sigma(\omega)$ peaks at 5.4×10^3 (xx , 6.21 eV) and 7.6×10^3 (zz , 6.43 eV); increases to 4.8×10^3 (xx) and 9.9×10^3 (zz) at $y = 0.50$; then drops to 7.1×10^3 (xx) and 5.2×10^3 (zz) at $y = 0.75$. Compared to the parent compound, both dopants enhance $\sigma(\omega)$ —more significantly along the zz direction—with the strongest photoconductivity observed at intermediate concentrations. The refractive index $n(\omega)$, shown in Fig. 7(e) and 8(e), reveals notable changes upon doping. The undoped $\text{LaBi}_2\text{ClO}_4$ has static refractive indices of 1.88 (xx) and 1.81 (zz). Upon Br doping, $n(0)$ increases gradually relative to the parent: from 1.91/1.84 at $y = 0.25$, to 1.94/1.87 at $y = 0.50$, and 1.98/1.89 at $y = 0.75$. In contrast, I doping leads to a more pronounced increase compared to the parent values, with $n(0)$ rising to 2.02/1.93 ($y = 0.25$), 2.48/2.05 ($y = 0.50$), and 2.89/2.13 ($y = 0.75$), indicating significantly enhanced optical confinement. The extinction coefficient $k(\omega)$, depicted in Fig. 7(f) and 8(f), reflects absorption and scattering behavior. For Br doping, $k(\omega)$ shows peaks above 6 eV across all doping levels, with slight redshifts as y increases. In contrast, I doping results in more pronounced shifts: 6.22 eV ($y = 0.25$), 5.18 eV ($y = 0.50$), and 4.37 eV ($y = 0.75$), consistent with the band gap reduction observed earlier. Fig. 7(g) and 8(g) display the absorption coefficient $\alpha(\omega)$. In $\text{LaBi}_2\text{ClO}_4$, absorption begins around 3.65 eV and rises steeply, with strong peaks up to 11.5 eV. Br doping causes a moderate redshift of the absorption edge across $y = 0.25$ to 0.75 , while I doping causes a significant redshift with sharper absorption features, especially at $y = 0.50$ and 0.75 , signifying stronger interband transitions and enhanced optical activity. Finally, the reflectivity

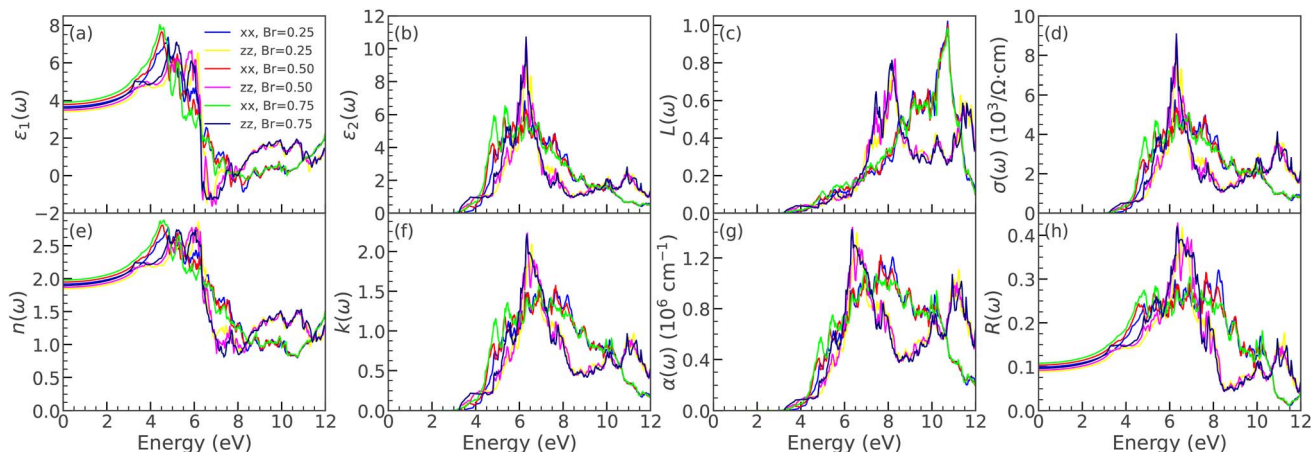


Fig. 7 Optical spectra of $\text{LaBi}_2\text{Cl}_{1-y}\text{Br}_y\text{O}_4$ ($y = 0.25, 0.50, 0.75$) using mBJ, showing (a)–(h): ϵ_1 , ϵ_2 , L , σ , n , k , α , and R for xx and zz polarizations.

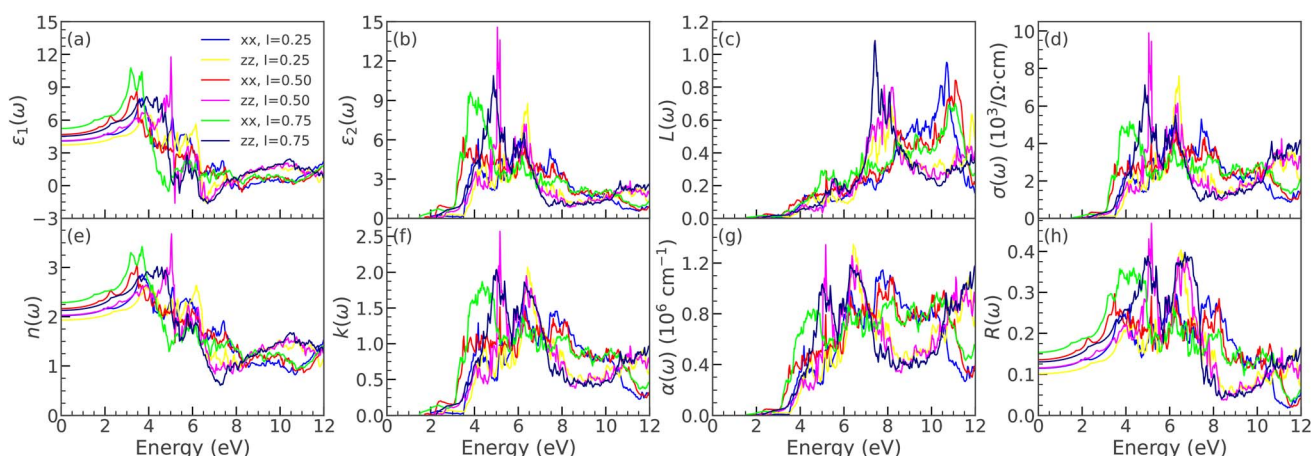


Fig. 8 (g) Optical spectra of $\text{LaBi}_2\text{Cl}_{1-y}\text{I}_y\text{O}_4$ ($y = 0.25, 0.50, 0.75$) using mBJ, showing (a)–(h): ϵ_1 , ϵ_2 , L , σ , n , k , α , and R for xx and zz polarizations.

spectra $R(\omega)$ (Fig. 7(h) and 8(h)) reveal moderate changes due to doping. For the parent compound, reflectivity peaks at 34% (xx) and 43% (zz). Br-doped systems show a slight reduction in the xx direction from 32% ($x = 0.25$) to 30% ($y = 0.75$). In I-doped systems, the reflectivity shows stronger modulation, with a notable maximum of 47% in the zz direction at $y = 0.50$, indicating tunability of surface optical properties by iodine substitution for potential optoelectronic applications. Given the prominent role of halogen substitution in modulating electronic properties, $\text{LaBi}_2\text{ClO}_4$ was studied within the broader class of layered bismuth oxyhalides. Experimental studies report indirect band gaps of 2.40 eV for $\text{Bi}_2\text{YO}_4\text{Cl}$, 2.21 eV for $\text{Bi}_2\text{EuO}_4\text{Cl}$, and 1.89 eV for $\text{Bi}_2\text{NdO}_4\text{Cl}$, with DFT closely matching the indirect nature, although GGA underestimates some gaps (e.g., 1.23 eV for $\text{Bi}_2\text{NdO}_4\text{Cl}$).^{32,73} $\text{Bi}_2\text{LaO}_4\text{Cl}$, synthesized *via* the flux method, exhibits experimental optical band gap of 2.79 eV, in good agreement with our DFT-mBJ calculations, compared to 2.50 eV for $\text{Bi}_2\text{YO}_4\text{Cl}$ and 3.16 eV for $\text{Bi}_3\text{O}_4\text{Cl}$.⁷⁴ To further validate these results, the calculated optical properties of $\text{LaBi}_2\text{ClO}_4$ were compared with experimental trends for bismuth oxyhalides (BiOX , $X = \text{Cl}, \text{Br}, \text{I}$).

Substituting Cl with Br or I narrows the band gap and shifts the absorption edge toward visible light, consistent with BiOX trends: BiOCl (3.44 eV), BiOBr (2.76 eV), BiOI (1.85 eV),⁷⁵ and $\text{BiOBr}_{1-x}\text{I}_x$ solid solutions (2.87–1.89 eV).⁷⁶ These effects arise from modified Bi–halogen hybridization and $\text{O } 2p \rightarrow \text{Bi } 6p$ transitions, corroborating experimental band edge shifts. The agreement confirms the reliability of our calculations and highlights halogen-doped $\text{LaBi}_2\text{ClO}_4$ as a promising visible-light optoelectronic and photocatalytic material.

4 Conclusions

First-principles DFT computations have been conducted to comprehensively examine the properties—including structural, mechanical, electronic, and optical aspects—of $\text{LaBi}_2\text{ClO}_4$ and its halogen-substituted forms $\text{LaBi}_2\text{Cl}_{1-y}\text{X}_y\text{O}_4$ ($X = \text{Br}, \text{I}$) at varying doping levels ($y = 0.25, 0.5, 0.75$). The negative formation energies confirm the thermodynamic stability of both the pristine and doped phases. Mechanical stability of the parent compound is validated by the satisfaction of the Born stability criteria. Analysis of the electronic structure indicates that



pristine $\text{LaBi}_2\text{ClO}_4$ behaves as an indirect band gap semiconductor, with its band gap gradually narrowing as halogen atoms are substituted. Notably, iodine doping induces a stronger band gap reduction than bromine, attributed to greater lattice distortion. Optical properties display pronounced anisotropy, with halogen doping enhancing dielectric constants and refractive indices, while shifting absorption edges toward lower energies. These tunable optoelectronic properties highlight the potential of $\text{LaBi}_2\text{ClO}_4$ and its halogen-doped derivatives for application in next-generation optoelectronic devices.

Author contributions

D. P. K. performed the DFT calculations focusing on the structural, electronic, and optical properties, and wrote the first draft of the manuscript in consultation with all co-authors. K. N. J. contributed to the DFT calculations related to thermodynamic stability and assisted in writing the manuscript, particularly the Introduction section. K. B. R. and S. L. carried out the calculations related to mechanical stability and properties. R. J. Y. and A. D. M. supervised part of the project, provide resources and analyzed the results. M. P. G. conceptualized and supervised the project, provided necessary resources, and directed the analysis in addition to writing the final version of the manuscript. All authors reviewed the content thoroughly and collectively endorsed the final manuscript for submission.

Conflicts of interest

The authors report no conflicts of interest.

Data availability

Most of the data that are used are kept within the manuscript and in the supplementary information (SI). The remaining data that support the findings of this research work will be made available from the corresponding author upon reasonable request. Supplementary information: formation energy, optimization, electronic band structures/DOS and related properties of $\text{LaBi}_2\text{Cl}_{1-y}\text{X}_y\text{O}_4$. See DOI: <https://doi.org/10.1039/d5ra06049d>.

Acknowledgements

This work was supported by a grant from University Grants Commission, Nepal from the Collaborative Research Grants (award no. CRG-78/79 S&T-03). Part of this work is also supported by National Priority Collaborative Research Grants (award no. EF-80/81-Col-9) from Research Directorate, Office of Rector, Tribhuvan University, Nepal. D. P. K. gratefully acknowledges the Research Directorate, Tribhuvan University, for the Master's Thesis Award that supported this work. M. P. G. acknowledges IFW-Dresden for providing the large-scale compute nodes to Tribhuvan University for scientific computations. M. P. G. and S. L. thanks Ulrike Nitzsche for technical assistance.

Notes and references

- 1 X. Chen, J. Zhao, G. Li, D. Zhang and H. Li, *Energy Mater.*, 2022, **2**, 200001.
- 2 J. Z. Hassan, A. Raza, U. Kumar and G. Li, *Sustain. Mater. Technol.*, 2022, **33**, e00478.
- 3 A. J. Yang, S. X. Wang, J. Xu, X. J. Loh, Q. Zhu and X. R. Wang, *ACS Nano*, 2023, **17**, 9748–9762.
- 4 P. Joshi, G. B. Acharya, I. B. Khadka, B. Srinivasan, S.-H. Kim and M. P. Ghimire, *ACS Appl. Energy Mater.*, 2025, **8**, 3217–3228.
- 5 S. D. Richardson and S. Y. Kimura, *Anal. Chem.*, 2019, **92**, 473–505.
- 6 H. Wu, L. Li, S. Wang, N. Zhu, Z. Li, L. Zhao and Y. Wang, *Phys. Chem. Chem. Phys.*, 2023, **25**, 25899–25924.
- 7 X. Wei, M. U. Akbar, A. Raza and G. Li, *Nanoscale Adv.*, 2021, **3**, 3353–3372.
- 8 L. Li, P. A. Salvador and G. S. Rohrer, *Nanoscale*, 2014, **6**, 24–42.
- 9 D. Kato, H. Suzuki, R. Abe and H. Kageyama, *Chem. Sci.*, 2024, **15**, 11719–11736.
- 10 X. Zhang, Z. Ai, F. Jia and L. Zhang, *J. Phys. Chem. C*, 2008, **112**, 747–753.
- 11 J. Li, Y. Yu and L. Zhang, *Nanoscale*, 2014, **6**, 8473–8488.
- 12 D. Huang, Z. Li, G. Zeng, C. Zhou, W. Xue, X. Gong, X. Yan, S. Chen, W. Wang and M. Cheng, *Appl. Catal., B*, 2019, **240**, 153–173.
- 13 X.-Y. Zhang, J.-J. Wang, P. Li, Z.-Y. Tan, J.-H. Zeng, Y.-R. He and N. Habibul, *Chem. Eng. J.*, 2022, **428**, 131209.
- 14 A. Fujishima and K. Honda, *Nature*, 1972, **238**, 37–38.
- 15 F. Irfan, M. U. Tanveer, M. A. Moiz, S. W. Husain and M. Ramzan, *Eur. Phys. J. B*, 2022, **95**, 184.
- 16 K. M. Lee, C. W. Lai, K. S. Ngai and J. C. Juan, *Water Res.*, 2016, **88**, 428–448.
- 17 S. G. Botta, J. A. Navio, M. C. Hidalgo, G. M. Restrepo and M. I. Litter, *J. Photochem. Photobiol., A*, 1999, **129**, 89–99.
- 18 C. Sun, J. Yang, M. Xu, Y. Cui, W. Ren, J. Zhang, H. Zhao and B. Liang, *Chem. Eng. J.*, 2022, **427**, 131564.
- 19 T. K. Ghorai, M. Chakraborty and P. Pramanik, *J. Alloys Compd.*, 2011, **509**, 8158–8164.
- 20 J. Choi, J. Lee, T. Lee and J. Korean Inst, *Electr. Electron. Mater. Eng.*, 2024, **37**, 231–240.
- 21 J. Schneider, M. Matsuoka, M. Takeuchi, J. Zhang, Y. Horiuchi, M. Anpo and D. W. Bahnemann, *Chem. Rev.*, 2014, **114**, 9919–9986.
- 22 A. Sudhaik, P. Raizada, P. Shandilya, D.-Y. Jeong, J.-H. Lim and P. Singh, *J. Ind. Eng. Chem.*, 2018, **67**, 28–51.
- 23 S. O. Ganiyu, C. A. Martinez-Huitle and M. A. Rodrigo, *Appl. Catal., B*, 2020, **270**, 118857.
- 24 L. Zhu, Y. Xie, X. Zheng, X. Yin and X. Tian, *Inorg. Chem.*, 2002, **41**, 4560–4566.
- 25 B. Zhang, G. Ji, Y. Liu, M. Gondal and X. Chang, *Catal. Commun.*, 2013, **36**, 25–30.
- 26 Z. Wang, M. Chen, D. Huang, G. Zeng, P. Xu, C. Zhou, C. Lai, H. Wang, M. Cheng and W. Wang, *Chem. Eng. J.*, 2019, **374**, 1025–1045.



- 27 M. Schmidt, H. Oppermann, C. Hennig, R. W. Henn, E. Gmelin, N. Soger and M. Binnewies, *Z. Anorg. Allg. Chem.*, 2000, **626**, 125–135.
- 28 U. Eggenweiler, E. Keller, V. Krämer, C. Meyer and J. Ketterer, *Z. Kristallogr. New Cryst. Struct.*, 1998, **213**, 735.
- 29 C. J. Milne, P. Lightfoot, J. D. Jorgensen and S. Short, *J. Mater. Chem.*, 1995, **5**, 1419–1421.
- 30 F. Deng, Y. Luo, H. Li, B. Xia, X. Luo, S. Luo and D. D. Dionysiou, *J. Hazard. Mater.*, 2020, **383**, 121127.
- 31 K. Ren, J. Liu, J. Liang, K. Zhang, X. Zheng, H. Luo, Y. Huang, P. Liu and X. Yu, *Dalton Trans.*, 2013, **42**, 9706–9712.
- 32 J. Pandey, N. Yadav, P. Yadav, S. Rao, K. Malla, P. Koirala, M. P. Ghimire and R. Nagarajan, *Inorg. Chem.*, 2023, **62**, 9471–9483.
- 33 N. Yadav and R. Nagarajan, *Solid State Sci.*, 2023, **139**, 107154.
- 34 K. Koepernik and H. Eschrig, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1999, **59**, 1743–1757.
- 35 P. Giannozzi, S. Baroni, N. Bonini, M. Calandra, R. Car, C. Cavazzoni, D. Ceresoli, G. L. Chiarotti, M. Cococcioni, I. Dabo, A. D. Corso, S. de Gironcoli, S. Fabris, G. Fratesi, R. Gebauer, U. Gerstmann, C. Gougoussis, A. Kokalj, M. Lazzeri, L. Martin-Samos, N. Marzari, F. Mauri, R. Mazzarello, S. Paolini, A. Pasquarello, L. Paulatto, C. Sbraccia, S. Scandolo, G. Sclauzero, A. P. Seitsonen, A. Smogunov, P. Umari and R. M. Wentzcovitch, *J. Phys.: Condens. Matter*, 2009, **21**, 395502.
- 36 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865–3868.
- 37 F. Tran and P. Blaha, *Phys. Rev. Lett.*, 2009, **102**, 226401.
- 38 J. Choi, J. Lee, T. Lee and J. Korean Inst, *Electr. Electron. Mater. Eng.*, 2024, **37**, 231–240.
- 39 H. Ko, Y.-w. Lim, S. Han, C. K. Jeong and S. B. Cho, *ACS Energy Lett.*, 2021, **6**, 2792–2799.
- 40 C. Malica and A. Dal Corso, *J. Phys. Condens. Matter*, 2020, **32**, 315902.
- 41 F. D. Murnaghan, *Proc. Natl. Acad. Sci. U. S. A.*, 1944, **30**, 244–247.
- 42 J. E. Saal, S. Kirklin, M. Aykol, B. Meredig and C. Wolverton, *Jom*, 2013, **65**, 1501–1509.
- 43 Q. Peng, A. R. Zamiri and S. De, *arXiv*, 2011, preprint, arXiv:1107.1453, DOI: [10.48550/arXiv.1107.1453](https://doi.org/10.48550/arXiv.1107.1453).
- 44 F. Mouhat and F. m. c.-X. Coudert, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2014, **90**, 224104.
- 45 G. D. Zartman, H. Liu, B. Akdim, R. Pachter and H. Heinz, *J. Phys. Chem. C*, 2010, **114**, 1763–1772.
- 46 R. Yu, X. Chong, Y. Jiang, R. Zhou, W. Yuan and J. Feng, *RSC Adv.*, 2015, **5**, 1620–1627.
- 47 D. Chattaraj, C. Majumder and S. Dash, *J. Alloys Compd.*, 2014, **615**, 234–242.
- 48 S. F. Pugh, *Philos. Mag.*, 1954, **45**, 823–843.
- 49 J. Haines, J. M. Leger and G. Bocquillon, *Annu. Rev. Mater. Res.*, 2001, **31**, 1–23.
- 50 R. K. Joshi, S. R. Bhandari and M. P. Ghimire, *RSC Adv.*, 2022, **12**, 24156–24162.
- 51 J. Camargo-Martínez and R. Baquero, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2012, **86**, 195106.
- 52 D. R. Jaishi, N. Sharma, B. Karki, B. P. Belbase, R. P. Adhikari and M. P. Ghimire, *AIP Adv.*, 2021, **11**, 025304.
- 53 M. P. Ghimire, C. Sandeep, T. P. Sinha and R. K. Thapa, *Phys. B: Condens. Matter*, 2011, **406**, 3454–3457.
- 54 Y. Ning, M. Wang, X. Jia, H. Ye, J. Yu and Y. Wu, *J. Phys. Chem. C*, 2025, **129**, 7040–7051.
- 55 K.-L. Zhang, C.-M. Liu, F.-Q. Huang, C. Zheng and W.-D. Wang, *Appl. Catal., B*, 2006, **68**, 125–129.
- 56 Z.-Y. Zhao and W.-W. Dai, *Inorg. Chem.*, 2014, **53**, 13001–13011.
- 57 T. Das and S. Datta, *Nanoscale Adv.*, 2020, **2**, 1090–1104.
- 58 Y. Wang, J. Meng, S. Jing, K. Wang, C. Ban, Y. Feng, Y. Duan, J. Ma, L. Gan and X. Zhou, *Energy Environ. Mater.*, 2023, **6**, e12432.
- 59 A. M. Ganose, M. Cuff, K. T. Butler, A. Walsh and D. O. Scanlon, *Chem. Mater.*, 2016, **28**, 1980–1984.
- 60 B. P. Belbase, D. R. Jaishi, R. P. Adhikari and M. P. Ghimire, *Solid State Commun.*, 2023, **361**, 115085.
- 61 M. N. Islam and J. Podder, *Mater. Sci. Semicond. Process.*, 2021, **121**, 105419.
- 62 C. Zhang, Y. Jia, Y. Jing and Y. Yao, *Int. J. Hydrogen Energy*, 2015, **40**, 1343–1351.
- 63 K. Ren, J. Yu and W. Tang, *J. Alloys Compd.*, 2020, **812**, 152049.
- 64 X. Yang, X. Qin, J. Luo, N. Abbas, J. Tang, Y. Li and K. Gu, *RSC Adv.*, 2020, **10**, 2615–2623.
- 65 D. P. Rai, A. Shankar, M. P. Ghimire, C. Sandeep and R. K. Thapa, *Comput. Mater. Sci.*, 2015, **101**, 313–320.
- 66 G. Bahadur Acharya, S.-H. Kim and M. P. Ghimire, *Modell. Simul. Mater. Sci. Eng.*, 2024, **32**, 065012.
- 67 G. B. Acharya, B. P. Belbase and M. P. Ghimire, *Electron. Struct.*, 2023, **5**, 045008.
- 68 D. Abdullah and D. C. Gupta, *Sci. Rep.*, 2024, **14**, 23575.
- 69 H. He, R. Orlando, M. A. Blanco, R. Pandey, E. Amzallag, I. Baraille and M. Rérat, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2006, **74**, 195123.
- 70 Y. Liu, Q. Hou, S. Sha and Z. Xu, *Vacuum*, 2020, **173**, 109127.
- 71 T. Y. Ahmed, S. B. Aziz and E. M. Dannoun, *Heliyon*, 2024, **10**, year.
- 72 M. Zaniib, M. W. Iqbal, M. Manzoor, M. Asghar, R. Sharma, N. N. Ahmad, S. M. Wabaidur, M. A. Habila, S. A. Abdelmohsen, A. M. Abdelbacki, *et al.*, *J. Mater. Sci. Eng. B*, 2023, **295**, 116604.
- 73 X. Wang, J. Ran, M. Tao, Y. He, Y. Zhang, X. Li and H. Huang, *Mater. Sci. Semicond. Process.*, 2016, **41**, 317–322.
- 74 A. Nakada, D. Kato, R. Nelson, H. Takahira, M. Yabuuchi, M. Higashi, H. Suzuki, M. Kirsanova, N. Kakudou, C. Tassel, T. Yamamoto, C. M. Brown, R. Dronskowski, A. Saeki, A. Abakumov, H. Kageyama and R. Abe, *J. Am. Chem. Soc.*, 2021, **143**, 2491–2499.
- 75 H. An, Y. Du, T. Wang, C. Wang, W. Hao and J. Zhang, *Rare Met.*, 2008, **27**, 243–250.
- 76 X. Zhang, C.-Y. Wang, L.-W. Wang, G.-X. Huang, W.-K. Wang and H.-Q. Yu, *Sci. Rep.*, 2016, **6**, 22800.

